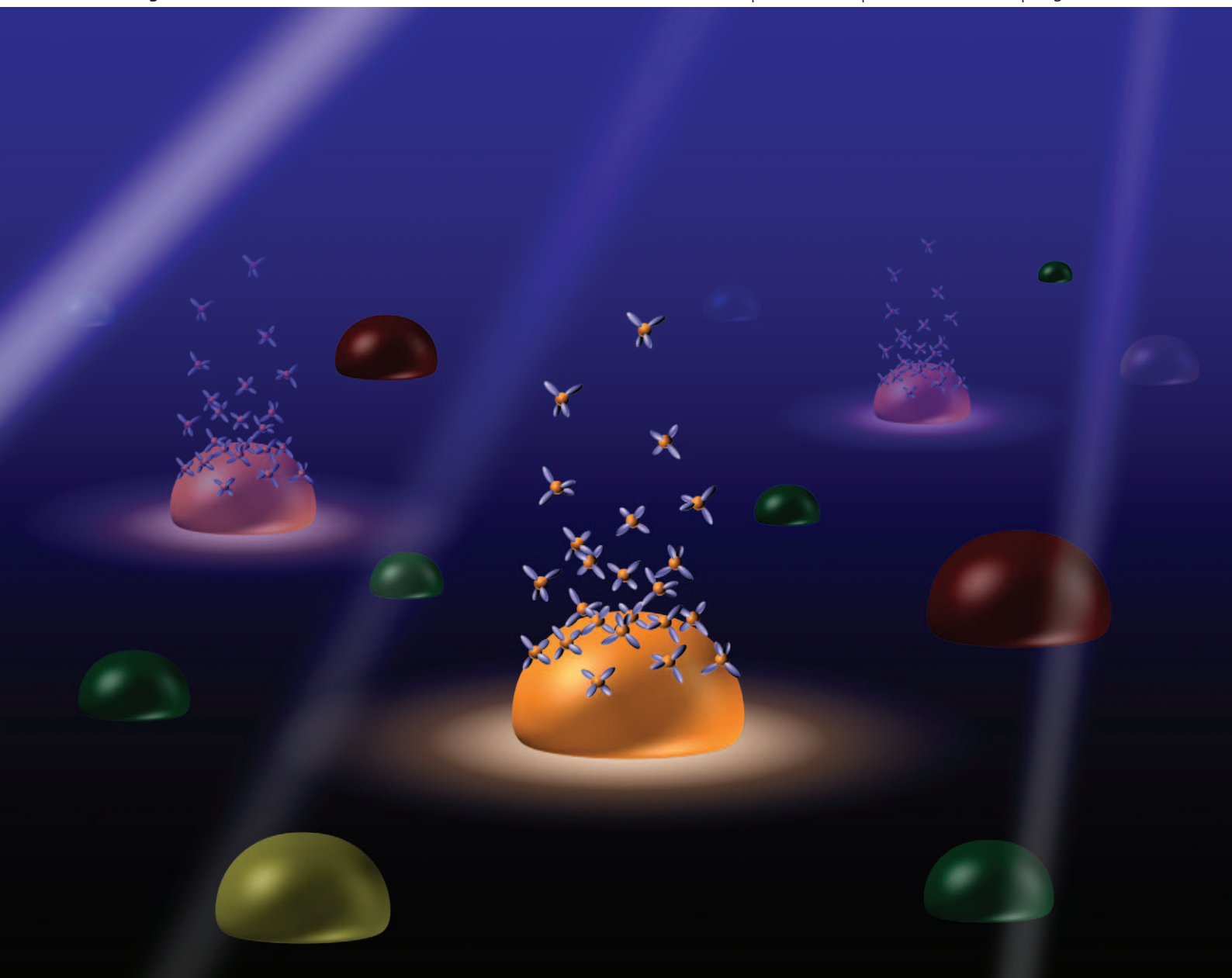


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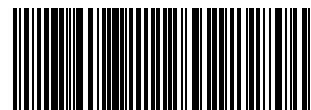
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**COVER ARTICLE**

Tatsuma *et al.*

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# Plasmon-induced oxidation of gold nanoparticles on TiO<sub>2</sub> in the presence of ligands

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Yohei Konishi, Ichiro Tanabe† and Tetsu Tatsuma\*

The mechanism of plasmon-induced charge separation of gold nanoparticles (Au NPs) on TiO<sub>2</sub> was investigated with the help of anionic ligands. Au NPs are oxidized in the presence of KX (X = SCN, Br or Cl), chiefly via direct oxidation to [AuX<sub>4</sub>]<sup>−</sup>. The reactivity of Au NPs depends on the solution pH, the type of anion ligands and the excitation wavelength, suggesting that the photopotential of Au NPs during plasmon-induced charge separation depends on both the flatband potential of TiO<sub>2</sub> and the energy of irradiated photons. The results indicate that the reactivity and efficiency of the plasmon-induced charge separation and accompanying reactions can be tuned and optimized by changing those factors.

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## 1. Introduction

Noble metal nanoparticles (NPs) absorb and scatter light at specific wavelengths due to localized surface plasmon resonance (LSPR). As their resonance wavelength and intensity depend largely on the particle size, shape, interparticle distance and dielectric environment, they exhibit a variety of colours.<sup>1,2</sup> In our past research, we found that Au,<sup>3,4</sup> Ag<sup>4,5</sup> and Cu<sup>6</sup> NPs on TiO<sub>2</sub> or ZnO exhibit plasmon-induced charge separation, which is caused by light irradiation at LSPR wavelengths. Stable Au NPs have been applied to plasmon-based photocatalysis<sup>3,7,8</sup> and photovoltaic cells.<sup>3,4,9–11</sup> On the other hand, less stable Ag NPs undergo photo-oxidation and dissolution,<sup>12</sup> which have been applied to multicolour,<sup>5,13</sup> infrared<sup>14</sup> and high speed<sup>15</sup> photochromism, as well as photoinduced multicolour control of single NP<sup>16</sup> and photomorphing gels.<sup>17</sup> However, the mechanisms of the plasmon-induced charge separation have not yet been elucidated completely, and learning the details of the phenomenon may lead to an improvement of the reaction efficiency and the development of other applications.

We have reported that TiO<sub>2</sub> electrodes modified with Au<sup>3,4</sup> or Ag<sup>4,18</sup> NPs exhibit negative potential shifts and anodic currents under illumination. In contrast, Au<sup>9,19</sup> or Ag<sup>9,20</sup> NPs coated with TiO<sub>2</sub> exhibit positive potential shifts and cathodic currents. On the basis of these results, we concluded that the charge separation is caused by electron transfer from metal NPs to TiO<sub>2</sub>, which is also supported in related studies by other groups.<sup>7,8,10,11,21–23</sup>

As mentioned above, Au NPs are generally so stable that the plasmon-induced charge separation does not cause their oxidative dissolution.<sup>3</sup> In our previous research, however, we have achieved the plasmon-induced oxidative dissolution of Au NPs with the aid of I<sup>−</sup> (KI aqueous solution) and developed a multicolour change material, which is more resistive to white light than a Ag-based material.<sup>24</sup> However, it is yet to be elucidated whether the Au dissolution is predominantly due to direct oxidation (Au + 4I<sup>−</sup> → [AuI<sub>4</sub>]<sup>−</sup> + 3e<sup>−</sup>, E° = +0.56 V vs. NHE) or indirect oxidation via I<sub>3</sub><sup>−</sup> (3I<sup>−</sup> → I<sub>3</sub><sup>−</sup> + 2e<sup>−</sup>, E° = +0.55 V, then 2Au + 3I<sub>3</sub><sup>−</sup> → 2[AuI<sub>4</sub>]<sup>−</sup> + I<sup>−</sup>), because the potential of the former reaction is close to that of the latter (Fig. 1a). To obtain further information about this, here we used other ligands, *i.e.* Br<sup>−</sup>, Cl<sup>−</sup> and SCN<sup>−</sup>. In the case of Br<sup>−</sup> and Cl<sup>−</sup>, the potentials of the direct oxidation are far more negative than the oxidation potentials of the halide ion to tri-halide ion. Also, in a solution of SCN<sup>−</sup>, the indirect mechanism would not hold as SCN<sub>3</sub><sup>−</sup> is known to be unstable in aqueous media.<sup>25</sup> In addition, examination of the reactions in the presence of different ligands and at different pH values sheds light on the photopotentials of Au NPs during the plasmon-induced charge separation.

## 2. Experimental

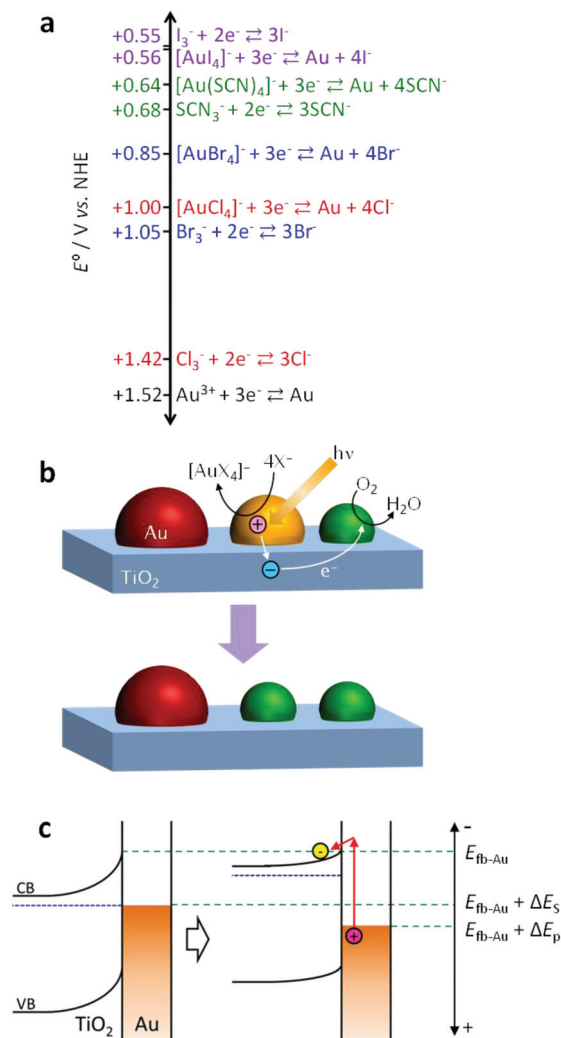
### Preparation of a TiO<sub>2</sub> film and Au NPs

A TiO<sub>2</sub> film loaded with Au NPs was prepared as previously reported.<sup>24</sup> A TiO<sub>2</sub> film was prepared on a Pyrex glass substrate from a commercial anatase sol (STS-21, Ishihara, 75% aq.) by a spin-coating method (at 1500 rpm for 10 s) and calcined at 723 K for 1 h. Au NPs were deposited photocatalytically on the TiO<sub>2</sub> film as follows. A mixture of 10 mM aqueous solution of HAuCl<sub>4</sub> and ethanol (1 : 9 by volume, 32 mL cm<sup>−2</sup>) was cast on the TiO<sub>2</sub> film and left for 15 min, followed by rinsing with water. The mixture was cast again and the film was irradiated

Institute of Industrial Science, University of Tokyo, 4-6-1 Komaba, Meguro-ku, Tokyo 153-8505, Japan. E-mail: tatsuma@iis.u-tokyo.ac.jp; Fax: +81-3-5452-6338; Tel: +81-3-5452-6336

† Present address: Department of Chemistry, School of Science and Technology, Kwansai Gakuin University, Sanda, Hyogo 669-1337, Japan.





**Fig. 1** (a) Values of the standard electrode potential  $E^\circ$  for oxidative coordination of Au with anionic ligands ( $X^-$ ) and oxidation of the anions to  $X_3^-$ . (b) Mechanism of spectral dip formation for the Au NP- $TiO_2$  system. (c) Potential shifts of Au NPs and  $TiO_2$  due to the plasmon-induced charge separation in an open circuit system.

with UV light (260–370 nm,  $25 \text{ mW cm}^{-2}$ ) for 10 s. Electrons in the  $TiO_2$  valence band (VB) are excited to the conduction band (CB), and used for the reduction of  $[AuCl_4]^-$  to Au NPs. The corresponding generated holes in the VB are consumed by oxidation of ethanol.

### Photoinduced reactions

Extinction (= absorption + scattering) spectra were collected using a V-670 spectrophotometer (Jasco). The as-prepared Au NP- $TiO_2$  sample was soaked in 0.5 M KSCN, KBr or KCl aqueous solution until the spectrum was stabilized. Then, the sample in the solution was irradiated with 600–700, 700–800 or 800–1000 nm light ( $10 \text{ mW cm}^{-2}$ ) for 30 min. The light source was a Xe lamp (Luminar Ace LA-410UV, Hayashi Watch Works) (600–700 nm) or a halogen lamp (HA-150UX, Myutron) (700–800 and 800–1000 nm) equipped with a long-pass filter (SCF-50S-60R, SCF-50S-70R or CLDM-50S, Sigma Koki) and a

short-pass or band-pass filter (FSWP-800-25.0M or SPF-1000-1.00, CVI Melles Griot; ITF-50S-83RT, Sigma Koki).

For monochromatic light irradiation ( $4 \text{ mW cm}^{-2}$ ) in 0.5 M KCl at pH 7, 8 or 9 (adjusted with KOH), a band-pass filter (full width at half maximum = 10 nm) for 700 nm (Asahi Spectra) or 800 or 900 nm (CVI Melles Griot) was used with the Xe lamp (700 nm) or the halogen lamp (800 and 900 nm).

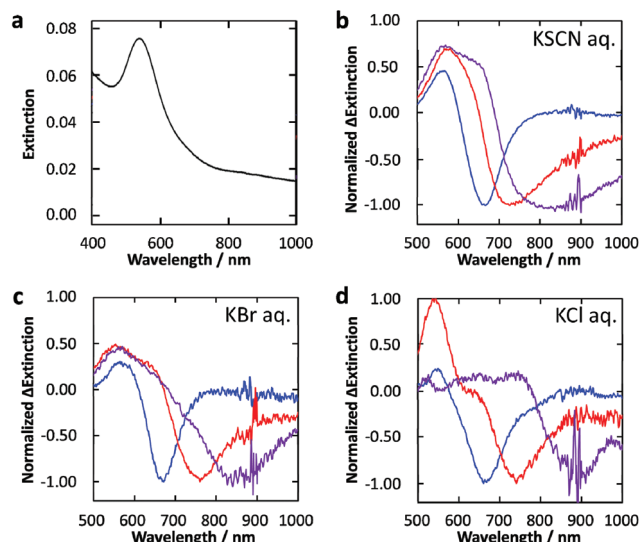
## 3. Results and discussion

### Spectral dip formation

The deposited Au NPs show an extinction peak at 545 nm and a tail extending to the near infrared region ( $\sim 1000 \text{ nm}$ ) (Fig. 2a) as a result of the formation of NPs with various size and anisotropy on the contact area with  $TiO_2$  which exhibit different LSPR wavelengths.<sup>24</sup>

The Au NP-modified  $TiO_2$  sample was irradiated with 600–700, 700–800 or 800–1000 nm light ( $10 \text{ mW cm}^{-2}$ ) in 0.5 M KSCN, KBr or KCl aqueous solution for 30 min. Normalized difference extinction spectra after the light irradiations are shown in Fig. 2b–d. In all these cases, extinction decreased preferentially at around the irradiation wavelengths and wavelength-selective extinction dips were formed. These spectral changes are explained in terms of selective oxidation and dissolution of resonant Au NPs to  $[AuX_4]^-$  (X: halogen or SCN) (Fig. 1b).<sup>24</sup> The peaks at  $\sim 540 \text{ nm}$  in the difference spectra are attributed to smaller NPs formed eventually by the dissolution, which exhibit LSPR at  $\sim 540 \text{ nm}$ .<sup>24</sup> Small Au NPs deposited by recombination of the  $[AuX_4]^-$  ions and electrons transferred to  $TiO_2$  may also be responsible for the peaks.<sup>24,26</sup>

In the case of X = SCN, the dip formation must be caused by the direct oxidation ( $Au + 4SCN^- \rightarrow [Au(SCN)_4]^- + 3e^-$ ),



**Fig. 2** (a) Extinction spectrum of Au NPs deposited photocatalytically on the  $TiO_2$  film. (b–d) Normalized difference extinction spectra of Au NPs on  $TiO_2$  after 600–700 (blue), 700–800 (red) and 800–1000 (purple) nm light irradiation ( $10 \text{ mW cm}^{-2}$ , 30 min) in 0.5 M aqueous solution of (b) KSCN, (c) KBr and (d) KCl.



because  $\text{SCN}_3^-$  is unstable in aqueous media,<sup>25</sup> as described above. The direct oxidation should also be the major pathway in the case of  $X = \text{Br}$  and  $\text{Cl}$ , because the potentials for the direct oxidation are more negative than those for the  $\text{X}_3^-$  generation (Fig. 1a). In particular, the potential for  $\text{Cl}_3^-$  generation is even more positive than that of water oxidation.

### Photoreactivity at different pH

Next we examined the dependence of the photoreactivity on the solution pH, which affects the flatband potential of  $\text{TiO}_2$  at the  $\text{TiO}_2$ -solution interface  $E_{\text{fb-sol}}$ . Actually, the flatband potential of  $\text{TiO}_2$  is known to shift by  $-0.06$  V per pH unit due to deprotonation of surface hydroxyl groups.<sup>27</sup> We irradiated the Au NP-modified  $\text{TiO}_2$  sample with monochromatic light at 700, 800 and 900 nm ( $4 \text{ mW cm}^{-2}$ ,  $\text{fwhm} = 10 \text{ nm}$ ) in 0.5 M aqueous KCl at pH 7, 8 or 9 (adjusted with KOH; the film was damaged at pH 10). As a result, dip formation was observed at all the wavelengths at pH 7 (Fig. 3a), whereas it was not observed at 900 nm at pH 8 and 9 (Fig. 3b, c). The  $E_{\text{fb-sol}}$  values were estimated for the  $\text{TiO}_2$  film without Au NPs by voltammetry to be  $-0.14$ ,  $-0.20$  and  $-0.25$  V vs. NHE at pH 7, 8 and 9, respectively. On the other hand, the potential for the  $\text{Au}/[\text{AuCl}_4]^-$  redox couple  $E_{\text{Au-Cl}}$ , which does not involve  $\text{H}^+$  or  $\text{OH}^-$ , should not depend on pH.

There are two possible explanations for the suppression of the photoreactions at high pH. (I) If the  $\text{TiO}_2$  CB level at the  $\text{TiO}_2$ -Au NP interface  $E_{\text{fb-Au}}$  is not affected by the CB level at the  $\text{TiO}_2$ -solution interface  $E_{\text{fb-sol}}$ , the rise of the latter could slow down electron transport from the  $\text{TiO}_2$ -Au NP interface to other parts of  $\text{TiO}_2$  and thereby lower the charge separation efficiency. However, the dependence of the reactivity on the excitation wavelength only at higher pH is not explained by this hypothesis. (II) If  $E_{\text{fb-Au}}$  is affected by  $E_{\text{fb-sol}}$  because the area of the  $\text{TiO}_2$ -Au NP interface is so small, the

photopotential at Au NPs may depend on pH. The photopotential may also depend on the excitation wavelength.

In the plasmon-induced charge separation, some electrons in a resonant Au NP should be transferred to  $\text{TiO}_2$  through the Schottky barrier at the  $\text{TiO}_2$ -Au NP interface.<sup>3,19</sup> The barrier height  $\Delta E_{\text{S}}$  is estimated to be  $\sim 1.2$  eV or higher, as the electron affinity of nanoparticulate anatase  $\text{TiO}_2$  and the work function of Au are  $\sim 3.9$  eV<sup>28,29</sup> and  $5.1$ – $5.47$  eV,<sup>30,31</sup> respectively. Sönnichsen *et al.*<sup>32</sup> described that LSPR decays partially *via* inter- or intraband transition. However, it may not be reasonable to assume that those excited states are long-lived enough to be involved in the electron transfer from Au NPs to  $\text{TiO}_2$ , unless electrons pass through trap states, if any, at the  $\text{TiO}_2$ -Au interface. If not, it is likely that the electron transfer is due to hot electron injection or external photoelectric effect. Under irradiation of photons with an energy of  $\Delta E_{\text{p}}$  eV that is higher than the Schottky barrier height  $\Delta E_{\text{S}}$  (*i.e.*  $\Delta E_{\text{p}} \geq \Delta E_{\text{S}}$ ), those types of electron transfer may shift the potential of Au NPs positively up to  $E_{\text{fb-Au}} + \Delta E_{\text{p}}$  V (Fig. 1c). If this is the case, the spectral dip formation should be suppressed when the  $E_{\text{fb-Au}} + \Delta E_{\text{p}}$  is more negative than  $E_{\text{Au-Cl}} + \Delta E_{\text{op}}$ , where  $\Delta E_{\text{op}}$  is the overpotential needed for the redox reaction. That is, the photo-induced redox reaction could be suppressed at high pH and at a long wavelength even if  $\Delta E_{\text{p}} \geq \Delta E_{\text{S}}$ .

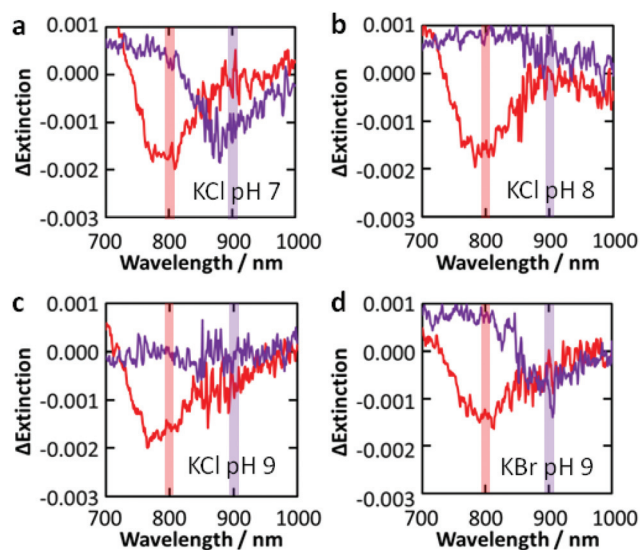
Actually, the reaction occurs under 700 and 800 nm light ( $\Delta E_{\text{p}} = 1.77$  and  $1.55$  eV, respectively) at pH 7–9, where  $E_{\text{fb-Au}} + \Delta E_{\text{p}} = +1.30$ – $+1.63$  V vs. NHE on the assumption that  $E_{\text{fb-Au}} = E_{\text{fb-sol}}$ . Also, it occurs under 900 nm light ( $\Delta E_{\text{p}} = 1.38$  eV) at pH 7 ( $E_{\text{fb-Au}} + \Delta E_{\text{p}} = +1.24$ ) but not at pH 8 and 9 ( $E_{\text{fb-Au}} + \Delta E_{\text{p}} = +1.13$ – $+1.18$ ). Therefore, the threshold exists at  $+1.18$ – $+1.24$  V. The difference between this threshold potential and the standard electrode potential for the formation of  $[\text{AuCl}_4]^-$  ( $+1.00$  V vs. NHE) may be explained in terms of the reaction overpotential and/or the difference between  $E_{\text{fb-Au}}$  and  $E_{\text{fb-sol}}$ .

### Photoreactivity in the presence of different anions

Replacement of KCl with KBr resulted in dip formation in all the cases examined, even at pH 9 under 900 nm light (Fig. 3d), where the  $E_{\text{fb-Au}} + \Delta E_{\text{p}}$  value is  $+1.13$  V vs. NHE. This value is sufficiently positive compared with the standard potential for the formation of  $[\text{AuBr}_4]^-$  ( $+0.85$  V vs. NHE; Fig. 1a), even taking the possible overpotential of  $0.18$ – $0.24$  V into account. It is thus suggested that the photopotential of Au NPs is dominated by  $E_{\text{fb}}$  and  $\Delta E_{\text{p}}$ . Since  $E_{\text{fb}}$  is related not only to pH but also to the coexisting cations and semiconductor, these factors as well as the light wavelength would be optimized for intended reactions.

## 4. Conclusions

Wavelength-selective spectral dip formation was observed for the Au NP- $\text{TiO}_2$  system in the presence of KX ( $X = \text{SCN}$ ,  $\text{Br}$  or  $\text{Cl}$ ). These results suggest that the dip formation is caused by direct oxidation of Au NPs to  $[\text{AuX}_4]^-$ . The reactivity of Au NPs



**Fig. 3** Difference extinction spectra of Au NPs on  $\text{TiO}_2$  after irradiation with 800 and 900 nm ( $4 \text{ mW cm}^{-2}$ , 30 min) light in 0.5 M KCl ((a) pH 7, (b) 8 and (c) 9) and (d) KBr (pH 9) aqueous solutions.



depends on both solution pH and excitation wavelength, suggesting that the photopotential of Au NPs depends on both the flatband potential of TiO<sub>2</sub> and the energy of irradiated photons. The results indicate that the reactivity and efficiency of the plasmon-induced charge separation can be tuned and optimized by changing those factors.

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## Notes and references

- S. Link and M. A. El-Sayed, *J. Phys. Chem. B*, 1999, **103**, 8410.
- T. Jensen, L. Kelly, A. Lazarides and G. C. Schatz, *J. Cluster Sci.*, 1999, **10**, 295.
- Y. Tian and T. Tatsuma, *J. Am. Chem. Soc.*, 2005, **127**, 7632.
- Y. Tian and T. Tatsuma, *Chem. Commun.*, 2004, 1810.
- Y. Ohko, T. Tatsuma, T. Fujii, K. Naoi, C. Niwa, Y. Kubota and A. Fujishima, *Nat. Mater.*, 2003, **2**, 29.
- T. Yamaguchi, E. Kazuma, N. Sakai and T. Tatsuma, *Chem. Lett.*, 2012, **41**, 1340.
- E. Kowalska, R. Abe and B. Ohtani, *Chem. Commun.*, 2009, 241.
- H. Kominami, A. Tanaka and K. Hashimoto, *Chem. Commun.*, 2010, **46**, 1287.
- Y. Takahashi and T. Tatsuma, *Appl. Phys. Lett.*, 2011, **99**, 182110.
- Z. H. Chen, Y. B. Tang, C. P. Liu, Y. H. Leung, G. D. Yuan, L. M. Chen, Y. Q. Wang, I. Bello, J. A. Zapien, W. J. Zhang, C. S. Lee and S. T. Lee, *J. Phys. Chem. C*, 2009, **113**, 13433.
- P. Reineck, G. P. Lee, D. Brick, M. Karg, P. Mulvaney and U. Bach, *Adv. Mater.*, 2012, **24**, 4750.
- T. Tatsuma, *Bull. Chem. Soc. Jpn.*, 2013, **86**, 1.
- K. Naoi, Y. Ohko and T. Tatsuma, *J. Am. Chem. Soc.*, 2004, **126**, 3664.
- E. Kazuma and T. Tatsuma, *Chem. Commun.*, 2012, **48**, 1733.
- N. Crespo-Monterio, N. Destouches, L. Bois, F. Chassagneux, S. Reynaud and T. Fournel, *Adv. Mater.*, 2010, **22**, 3166.
- I. Tanabe and T. Tatsuma, *Nano Lett.*, 2012, **12**, 5418.
- T. Tatsuma, K. Takada and T. Miyazaki, *Adv. Mater.*, 2007, **19**, 1249.
- K. Kawahara, K. Suzuki, Y. Ohko and T. Tatsuma, *Phys. Chem. Chem. Phys.*, 2005, **7**, 3851.
- N. Sakai, Y. Fujiwara, Y. Takahashi and T. Tatsuma, *ChemPhysChem*, 2009, **10**, 766.
- Y. Takahashi and T. Tatsuma, *Nanoscale*, 2010, **2**, 1494.
- Y. Nishijima, K. Ueno, Y. Yokota, K. Murakoshi and H. Misawa, *J. Phys. Chem. Lett.*, 2010, **1**, 2031.
- T. Tachikawa, T. Yonezawa and T. Majima, *ACS Nano*, 2013, **7**, 263.
- A. Furube, L. Du, K. Hara, R. Katoh and M. Tachiya, *J. Am. Chem. Soc.*, 2007, **129**, 14852.
- Y. Konishi, I. Tanabe and T. Tatsuma, *Chem. Commun.*, 2013, **49**, 606.
- O. Barbosa-Filho and A. J. Monhemius, *Trans. Inst. Min. Metall. C*, 1994, **103**, C105.
- K. Matsubara, K. L. Kelly, N. Sakai and T. Tatsuma, *Phys. Chem. Chem. Phys.*, 2008, **10**, 2263.
- A. Fujishima, A. Sakamoto and K. Honda, *Seisan Kenkyu*, 1969, **21**, 450.
- R. Könenkamp, *Phys. Rev. B: Condens. Matter*, 2000, **61**, 11057.
- F. Lenzmann, J. Krueger, S. Burnside, K. Brooks, M. Grätzel, D. Gal, S. Rühle and D. Cahen, *J. Phys. Chem. B*, 2001, **105**, 6347.
- H. B. Michaelson, *J. Appl. Phys.*, 1977, **48**, 4729.
- D. E. Eastman, *Phys. Rev. B: Solid State*, 1970, **2**, 1.
- C. Sönnichsen, T. Franzl, T. Wilk, G. von Plessen, J. Feldmann, O. Wilson and P. Mulvaney, *Phys. Rev. Lett.*, 2002, **88**, 077402.

