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Taming functionality: easy-to-handle chiral phosphiranes†

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Enantiopure chiral phosphiranes possessing a binaphthyl backbone demonstrate remarkable thermal stability, are highly resistant to air-oxidation and are effective ligands in catalytic asymmetric hydrosilylations.

The phosphiranes are heterocycles which possess a three-membered phosphorus-containing ring which is highly strained as a result of the small sum of bond angles at the phosphorus atom ($\sum(P) < 260^\circ$).¹ The resultant greater s-character in the pyramidalised structure is considered to manifest itself in the form of lowered HOMO and LUMO energy levels, conferring upon the ligand a lower σ -donor, but better π -acceptor electronic profile. Such characteristics are desirable because they offer the opportunity to synthesise ligand libraries for catalysis with a different window of activity from that offered by the more typical PR_3 reagents. These intriguing molecules are still overlooked because of their tendency to be unstable.^{1b,2} Phenylphosphirane **I** (Fig. 1) is stable at 0 °C for a month but undergoes decomposition at higher temperatures.^{3,4} Bulky substituents around the phosphorus lend stability to the ring; thus **II** is a distillable liquid⁵ and **III** is a colourless crystalline solid.⁶ The other strategy one can employ to stabilise a phosphirane (apart from metal coordination) is to incorporate the three-membered ring into an elaborate polycyclic structure. This niche approach was used to prepare the elegant BABAR-Phos **IV** series of phosphiranes which resist decomposition.⁷ Phosphirano-[1,2]-thiaphospholes **V** have [a]-annulated bicyclic structures, isomerising in their tungsten complexes over the temperature range 25–80 °C.⁸

We recently reported an electronic stabilising effect whereby increasing the amount of π -conjugation on the backbone of primary phosphines allowed us to prepare the novel air-stable, chiral derivatives (*S*)-**1a** and (*R*)-**1b**.⁹ We were keen therefore to use these precursors to prepare chiral phosphiranes, of which there are very few reports,¹⁰ and to examine their properties, structure and reactivity. Following the synthetic method of Kubiak and co-workers,¹¹ (*S*)-**1a** and (*R*)-**1b** were treated with two equivalents of methyllithium followed by

dichloroethane, and thus we were able to prepare the asymmetric compounds (*S*)-**2a** and (*R*)-**2b** in high yields (Scheme 1).

The ³¹P NMR spectroscopic data for the two phosphiranes are consistent with those found for other members of this class, showing a characteristic resonance at very high field ($\delta = -235.0$ ppm for (*S*)-**2a**, -235.4 for (*R*)-**2b**). The ¹H NMR spectrum shows four distinctive peaks for the protons on the heterocycle, with a complicated fine-structure arising from coupling to each other, coupling to the nearest aromatic proton, and coupling to the phosphorus nucleus. The two proton signals at lower field (*endo* protons) both show NOE correlations to the aromatic proton in the 3-position. Thus the phosphirane unit is rotating in solution, an observation further rationalised by quantum chemical calculations.† COSY NMR experiments revealed that the ²J-coupling of the phosphorus nucleus to the two *exo* protons is significantly larger (18.0/18.1 Hz in (*S*)-**2a**/*(R)*-**2b**) than the coupling to the *endo* protons (1.8/2.0 Hz in (*S*)-**2a**/*(R)*-**2b**). The latter smaller coupling was thereby not fully resolvable in the ³¹P NMR spectra. The ¹³C NMR spectrum also demonstrates the diastereotopic nature of the phosphirane-ring carbon atoms; two individual resonances are observed although the phosphorus–carbon coupling constants within the phosphirane ring are equivalent (¹J_{PC} = 40.3 Hz for (*S*)-**2a**; ¹J_{PC} = 40.1 Hz for (*R*)-**2b**).

Two important features of our novel phosphiranes lie in their high thermal stability and their excellent resistance to air-oxidation. Remarkably (*S*)-**2a** and (*R*)-**2b** showed no sign of decomposition when heated overnight in refluxing toluene (under nitrogen). Nor was there any evidence of decomposition when the compounds were left open to the atmosphere in chloroform solution for seven days. In fact, in the solid state they can be stored in air over several months without problems. Treatment of the phosphiranes by peroxides in reagent-grade tetrahydrofuran does oxidise the compounds, but this process takes place *via* a different mechanistic process than the oxidation by air.¹² A full rationale for this behaviour will be provided

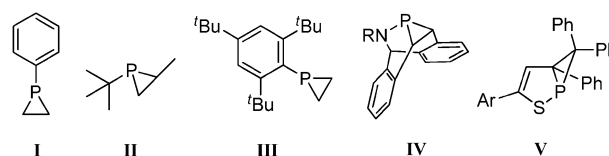
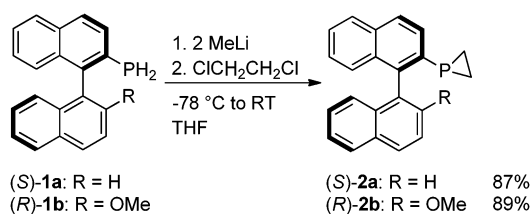


Fig. 1 Phosphiranes of varying thermal stability. **I** is unstable above RT; steric protection (**II**, **III**) and ring fusion aid stability (**IV**, **V**).

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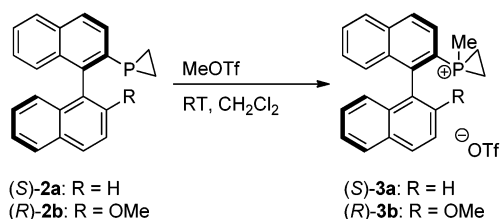


Scheme 1 Synthetic procedure for (*S*)-**2a** and (*R*)-**2b**.

elsewhere, as part of our larger study on the air-stability or otherwise of the phosphirane class of compounds in general.¹³

This remarkable behaviour led us to further investigate the reactivity of the phosphiranes. It has previously been reported that 2,6-dimethoxyphenylphosphirane reacts readily with sulfur at room temperature;¹⁴ (*S*)-**2a** and (*R*)-**2b** gave no reaction under these conditions, nor were any differences observed when the samples were heated at 50 °C in toluene for two hours. However, both (*S*)-**2a** and (*R*)-**2b** reacted with the strong methylating agent methyl trifluoromethanesulfonate which led to (*S*)-**3a** and (*R*)-**3b** in a quantitative conversion (Scheme 2). The same reactivity pattern was reported for 1-phenylphosphirane,¹⁵ while the stabilised ⁱPrBABARphos phosphirane ligand was found to be unreactive in this conversion,^{7a} further highlighting reactivity differences within the phosphirane class. The ³¹P NMR spectra showed single peaks at -101.2 ppm for (*S*)-**3a** and -102.0 ppm for (*R*)-**3b**, whilst in the ¹H NMR spectrum, the newly introduced methyl group appears as a doublet for both compounds at 1.60 ppm for (*S*)-**3a** (²J_{PH} = 17.9 Hz) and at 1.63 ppm for (*R*)-**3b** (²J_{PH} = 18.3 Hz).

In order to gain further insight into the geometry and electronic nature of the phosphirane group, we undertook a number of quantum chemical calculations of (*S*)-**2a** and (*R*)-**2b** at the B3LYP/6-31G* level of theory. The HOMO and LUMO energy levels of (*S*)-**2a** and (*R*)-**2b** were calculated and compared to their unstrained MOP derivatives, 2-(dimethylphosphino)-1,1'-binaphthyl **5a**† and 2-(dimethylphosphino)-2'-methoxy-1,1'-binaphthyl **5b**.† For each pair, introduction of the heterocycle lowers the energy of both the HOMO and LUMO. Thus we anticipate the phosphiranes to be poorer σ-donors but better π-acceptors than their corresponding unstrained MOP counterpart. Furthermore, the HOMO and LUMO energies for both H-MOP derivatives (*S*)-**2a** and **5a** were lower than their OMe-MOP counterparts, indicating that one has to consider not just the presence of the strained heterocycle but the backbone as a whole when ascertaining the relative energy levels of these compounds. Calculated bond angles around the phosphorus show a similar degree of pyramidalisation for (*S*)-**2a** and (*R*)-**2b** (Σ(P) 250/251°).



Scheme 2 Synthetic procedure for (*S*)-**3a** and (*R*)-**3b**.

We were keen to study the coordination chemistry of these binaphthyl-based phosphiranes with late transition metals as these are highly relevant for contemporary catalysis, but to date few examples have been reported.^{8a,10,11} Reaction of two equivalents of (*S*)-**2a** or (*R*)-**2b** with *cis*-[Pt(cod)Cl₂] resulted in the rapid, quantitative formation of **4a** or **4b** respectively as shown by HRMS, NMR, and X-ray crystallographic analyses. In the ³¹P NMR spectra of both complexes the ¹⁹⁵Pt satellites accompanying the central peak show large ¹J_{PtP} coupling constants (**4a**: -149.2 ppm, ¹J_{PtP} = 4170 Hz; **4b** -149.3 ppm, ¹J_{PtP} = 4160 Hz). This is a much larger coupling than those observed for the more conventional *cis*-[Pt(PR₃)₂Cl₂] complexes (¹J_{PtP} = 3400–3700 Hz),¹⁶ but relates well to that found for the only bisphosphirane platinum dichloride complex previously isolated (¹J_{PtP} = 4133 Hz).^{11b} The 10 Hz larger coupling for **4a** could be interpreted as a result of the ligand's higher π-acceptor strength in comparison to **4b**,^{6,17} which would also be in agreement with our calculated LUMO energy levels for these phosphirane ligands.†

Crystals suitable for X-ray analysis were grown from concentrated dichloromethane (**4a**, Fig. 2) or chloroform (**4b**, Fig. 3) solutions, and both compounds were obtained as solvates. Selected bond lengths and angles are given in the figure captions. Complex **4b** has crystallographic two-fold rotational symmetry, with the two phosphirane ligands equivalent, while **4a** has no imposed symmetry. The complexes each show a distorted square planar platinum(II) coordination geometry with typical bond lengths and angles about the platinum centre. The high degree of strain in the phosphirane ring is shown by the small C–P–C angles of 50.2(4) and 50.8(4)° for **4a**, and 50.9(3)° for **4b**. The heterocycles point over the aromatic rings of the second coordinated ligand (closest inter-ligand C⋯C distances: 3.5–3.7 Å). The steric demand of the methoxy group in **4b** appears responsible for the increased torsion angle of the binaphthyl backbone, the value of -94.7(8)° being significantly larger than those in

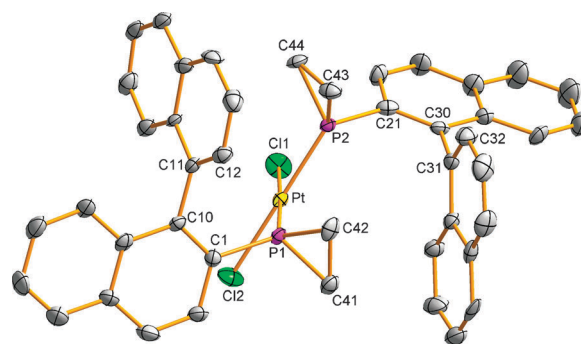


Fig. 2 View of the molecular structure of **4a** with 50% probability displacement ellipsoids. Hydrogen atoms have been omitted for clarity. Selected bond distances [Å] and angles [°]: Pt–Cl1 2.337(2), Pt–Cl2 2.334(2), Pt–P1 2.212(2), Pt–P2 2.209(2), P1–C41 1.802(3), P1–C42 1.816(8), P1–C1 1.814(7), C41–C42 1.534(11), P2–C43 1.786(8), P2–C44 1.782(7), P2–C21 1.794(8), C43–C44 1.529(11); C11–Pt–Cl2 90.52(8), C11–Pt–P2 87.06(10), Cl2–Pt–P1 85.40(9), P1–Pt–P2 97.02(8), Pt–P1–C1 118.8(2), C1–P1–C41 107.6(4), C1–P1–C42 111.7(4), C41–P1–C42 50.2(4), Pt–P2–C21 119.1(3), C21–P2–C43 109.8(4), C21–P2–C44 106.4(4), C43–P2–C44 50.8(4); C1–C10–C11–C12 -81.8(10), C21–C30–C31–C32 -88.5(9).

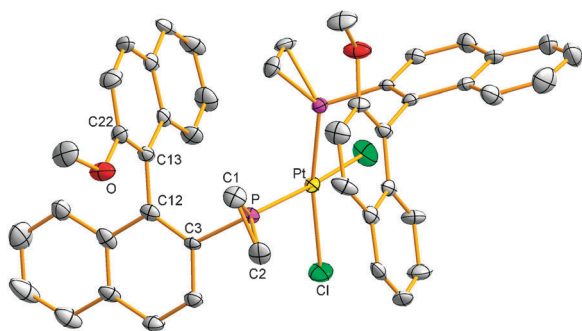
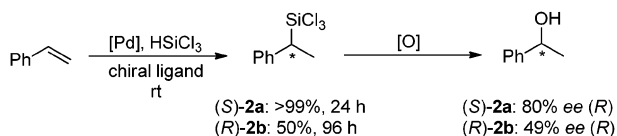


Fig. 3 View of the molecular structure of **4b**. Hydrogen atoms have been omitted for clarity. Selected average bond distances [Å] and angles [°]: Pt–Cl 2.338(3), Pt–P 2.204(3), P–C1 1.799(7), P–C2 1.811(7), P–C3 1.802(7), C1–C2 1.550(10); Cl–Pt–Cl' 90.42(15), Cl–Pt–P 87.28(7), P–Pt–P 96.92(14), Pt–P–C3 115.3(2), C1–P–C3 110.6(3), C2–P–C3 106.0(3), C1–P–C2 50.9(3); C3–C12–C13–C22 –94.7(8). A prime indicates a symmetry-equivalent atom.

4a of $-81.8(10)$ and $-88.5(9)^\circ$. These values compare well with the related complex *cis*-dichlorobis[1-(9-anthracene)phosphirane]platinum(II)^{11b} which exhibits slightly longer Pt–P bond lengths of 2.235(4) and 2.256(4) Å. This anthracenylphosphirane derivative has a slightly higher strain within its heterocyclic ring and is more pyramidalised at phosphorus (C–P–C bond angles: 49.7(8)–49.8(8)°) than **4a** or **4b**.

Having established that the thermal and air-stability of our phosphiranes does not infringe on their ability to coordinate to a late transition metal, we sought to extend the research to include a pilot study of the potential of these ligands for asymmetric catalysis. As a benchmark, the asymmetric induction capabilities of (*S*)-**2a** and (*R*)-**2b** were investigated in the catalytic hydrosilylation of styrene (Scheme 3).¹⁸ The reaction is required to be both highly regio- and enantioselective; oxidation of the initially formed trichlorosilane proceeds with retention of configuration and chiral 1-phenylethanol is produced (Scheme 3). Our catalysts were prepared *in situ* by reacting the chiral phosphirane ligand with [PdCl(η³-C₃H₅)₂].

The results obtained with the two ligands differ significantly. The hydrogen substituent in the 2'-position of the binaphthyl backbone seems to promote the activity as well as the enantioselectivity of the catalyst. A similar trend was found for the related parent MOP ligands when used in this transformation.¹⁹ The full conversion obtained with (*S*)-**2a** coupled with attaining an enantiomeric excess of 80% indicates the potential applications of this neglected class of ligand. To the best of our knowledge this is the first time that phosphiranes with the chirality located solely on the ligand backbone have



Scheme 3 The asymmetric hydrosilylation of styrene gives the chiral trichlorosilane, subsequent oxidation of which gives 1-phenylethanol.

been utilised in asymmetric catalysis.¹⁰ Note these reaction conditions are not optimised and the results represent only a first glance at the catalytic potential of these ligands.

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