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Synthesis of benzo[a]carbazole derivatives *via* intramolecular cyclization using Brønsted acidic carbonaceous material as the catalyst†

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In this work, a new procedure for the synthesis of benzo[a]carbazole from 1,3-diketones, primary amines, phenylglyoxal monohydrate, and malononitrile employing a solid acidic catalyst has been developed. The multicomponent reaction provided 3-cyanoacetamide pyrrole as an intermediate and then the formation of benzo[a]carbazole *via* intramolecular ring closure. The reaction was carried out for 2 h at 240 °C, resulting in the desired product with 73% yield. Acidic sites on the solid acid catalyst, made from rice husk-derived amorphous carbon with a sulfonic acid core (AC-SO₃H), provided the best activity. Acidic sites on the surface of the catalyst, including carboxylic, phenolic, and sulfonic acids, were 4.606 mmol g⁻¹ of the total acidity. AC-SO₃H demonstrated low cost, low toxicity, porosity, stability, and flexibility of tuning and reusability.

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Introduction

Carbazole frames are a peculiar type of N-heterocyclic compound. These compounds typically appear as structural motifs in both naturally occurring alkaloids and other synthetic chemicals. The physical properties of carbazoles are similar to those of synthetic dyes,¹ conducting polymers,² and optoelectronic components.³ Recently, several polymer derivatives have been incorporated into solar cells made of polymers (2,7-carbazole). Additionally, these compounds are frequently used as white, red, and green emitters in organic light-emitting diodes. By changing the structural characteristics of carbazoles at the C-2, -3, -6, -7, and -9 positions, it is possible to change their molecular and optical properties.⁴ Inorganic light-emitting systems have utilized a few benzo[a]carbazole, benzo[c] carbazole, and indolo[3,2-*b*]carbazole frames as molecular platforms, hole transport, and host materials. Recent research has revealed the usefulness of carbazole-containing ligands as anion receptors. For instance, derivatives of 3,6-dichlorocarbazole-1,8-diamide have the ability to hydrogen bond with anions. Notably, carbazoles are a diverse group of organic plant chemicals. Many of them possess strong biological properties, such as anticancer,⁵ antiepileptic,⁶ antibacterial,⁷ anti-inflammatory,⁸ antioxidative,⁹ analgesic,¹⁰

antidiarrheal,¹¹ antihistaminic,¹² neuroprotective,¹³ and pancreatic lipase inhibitory (Scheme 1).¹⁴

Recently, Knölker *et al.* thoroughly investigated the production of these beneficial carbazole alkaloids.¹⁵ Historically, carbazoles were made *via* processes like insertion of nitrene,¹⁶ indolization of Fischer,¹⁷ cyclization by Pummerer,¹⁸ Diels-Alder technique,¹⁹ cyclization *via* dehydrogenation of diaryl-amines.²⁰ Studies in more recent years have concentrated on transition metal-mediated processes, including ring-closing metathesis,²¹ cyclotrimerization,²² benzannulation,²³ Suzuki-Miyaura coupling,²⁴ and the creation of C-C and C-N bonds.^{25–27} To address issues with regiochemical selectivity and efficacy, many of these methods were continuously refined. This article's discussion of the most recent synthetic methods was organized according to how they create rings through the use of strategic bonds.

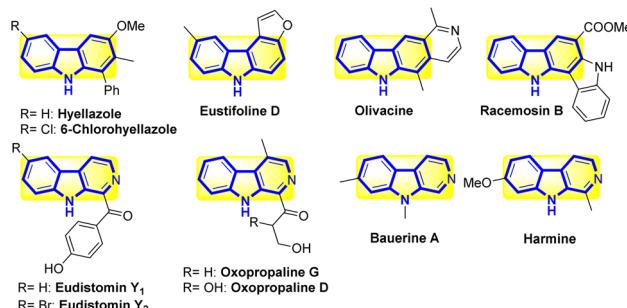
Maji *et al.* used Brønsted acid/base catalyzed for one-pot sequential-triple-relay benzannulation of 2-alkenyl indoles with low-cost aldehydes, a wide variety of structurally distinct carbazoles can be synthesized.²³ Deng *et al.* developed a straightforward three-component approach to produce diversely doped carbazoles through thermodynamically induced [2 + 2 + 2] annulation of indole, ketone, and nitro-alkene.²⁸ Liu *et al.* proposed a metal-free, sequential triple C-C coupling-radical-cyclization procedure to regioselectively generate substituted carbazole compounds using DMAP as the only reagent.²⁹ Through the transition-metal and exogenous-oxidant-free C-H bond amination approach, P. Zhang *et al.* also produced the carbazole-derived alkaloids clausine and glycozoline.³⁰ They also developed a technique for an environmentally secure electrochemical C-H bond dehydrogenative

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Scheme 1 Chemical structure of benzo[a]carbazole frames and their bioactive.

amination. This technique may be used to create preferred carbazole moieties, which were advantageous in a variety of situations. It was emphasized that this method's scalability, relatively mild conditions, and universal applicability made it more environmentally friendly and sustainable. A general technique for creating the carbazol-4-amine motif had been devised by D. Cao *et al.* by a vinylogous Michael addition, cyclization, isomerization, and elimination reaction of 3-nitroindoles with alkylidene malononitriles. The process of this method did not involve the usage of transition metals at any point.³¹

An innovative kind of solid Brønsted acid catalyst was amorphous carbon-carrying sulfonic acid groups. In recent years, organic synthesis, adsorbents, and electrochemistry had all researched novel forms of carbon materials in detail. Using readily accessible agricultural by-products, it was inexpensive and simple to create the sulfonated amorphous carbon, which has special physical and chemical characteristics including strong Brønsted acid, large surface area, high stability, reactivity, and cyclability. By incompletely carbonizing sulfopoly cyclic aromatic compounds in concentrated H_2SO_4 or sulfonating incompletely carbonized natural organic materials, carbon-based solid acids can be made easily. Three distinct functional groups including $-SO_3H$, $-COOH$, and $-OH$ groups, which containing in the material with carbon frames.³²⁻³⁴ The usual solid acids tested; however, all have a single functional group. The strong catalytic activity of the carbon material may be due to this. After the reaction, the carbon catalyst can be easily removed from the saccharide solution and reused without losing any activity.

Nowadays, the synthesis of carbonaceous framework materials from agricultural waste sources has received much attention and popular research around the world. Because of its exceptional characteristics, amorphous carbon materials are now the subject of a significant amount of study all around the globe in a variety of applications. These applications include biomedical, energy systems, and environmental remediation. To possibly produce amorphous carbon compounds, carbonaceous feedstocks such as rice byproducts (straw and husk), wood fiber, sugarcane bagasse, wood, sawdust, *etc.*, may be used.³⁵⁻³⁷

In our previous work, effective sulfonated amorphous carbon from rice husk carbonation is reported. By utilizing sulfuric acid for sulfonation, the AC- SO_3H was created. The AC- SO_3H catalyst activity was evaluated through the synthesis of benzo[a]carbazoles from 3-cyanoacetamide pyrroles through intramolecular cyclization. This method was considered a novel and effective method for synthesizing benzo[a]carbazole frames through intramolecular cyclization. Additionally, it attempted to explore the recycling of catalysts.

Experimental

Chemical and analytical techniques

All chemicals with high purity were purchased from Sigma-Aldrich and Acros. Analytical thin-layer chromatography (TLC) results were gathered on Merck F-254 silica gel-coated aluminum plates. Column chromatography using Merck silica gel (60, 230–400 mesh) was done. With the use of a Bruker Avance 500 MHz, the 1H and ^{13}C NMR spectra were recorded. Internal standards were either TMS or solvent peaks, and the solvent was DMSO-d6. Determining melting point included using the Buchi melting point B-545. ATR-FTIR spectra were sustained using a Bruker E400 FT-IR spectrometer. ATR-FTIR spectra in the region of 4000–600 cm^{-1} . TGA was measured using the Q-500 thermal gravimetric analyzer under airflow with a temperature gradient 5 $^{\circ}C$ min^{-1} . Powder X-ray diffraction (P-XRD) data for refinement were collected on a Bruker D8 Advance utilizing Ni-filtered Cu K ($\lambda = 1.54059$) radiation. Using the Hitachi S-4800 scanning electron microscope and the XZS-107T digital microscope connected to NHV-CAM through the program eScope, the materials' morphology was examined (SEM). The Quantachrome NOVA 3200e system was used to quantify the N_2 isotherm at 77 K. To ascertain the elemental composition of sorbents, energy-dispersive X-ray spectroscopy (EDX) examination was realized utilizing an EMAX energy EX-400 EDX instrument. The Bruker micrOTOF-QII MS operating at 80 eV was used to collect data for HRMS (ESI).

Typical procedure for the synthesis of amorphous carbon immobilized sulfonic acid group (AC- SO_3H)

Under N_2 atmosphere, at a temperature of 400 $^{\circ}C$ prolong 10 hours, rice husks were partly carbonized. After heating the combination for ten hours at 150 $^{\circ}C$ while it was being purged with nitrogen, powerful sulfuric acid was incorporated to the mixture. The dark material was washed many times with deionized water hot at 80 $^{\circ}C$ until the presence of sulfate anions was no longer detectable in the filtrate. The sulfonated amorphous carbon catalyst subsequently went through a drying process that lasted for two hours and was carried out at a temperature of 100 $^{\circ}C$ in air. FT-IR, XRD, TGA, EDS, and SEM were the instruments that were used, the catalyst's purity and structural integrity were verified.

General procedure for measuring the acidity of the AC- SO_3H catalyst

0.25 g of AC or AC- SO_3H was added to a sodium hydroxide aqueous solution (0.05 N, 30 mL) (NaOH was calibrated with



oxalic acid). At room temperature, the mixture was stirred for 60 min. Following centrifugal separation, hydrochloric acid (0.05 N) aqueous solution was used to titrate the supernatant solution using phenolphthalein as an indicator. Record the pH of the solution every 0.1 mL of 0.05 N HCl.

Methodological outline for the synthesis of benzo[*a*]carbazoles

In a 25 mL flask, a mixture of 3-cyanoacetamide pyrrole scaffolds (0.25 mmol) and AC-SO₃H (6 mg) in the presence of DMSO (5 mL). The reaction was stirred and refluxed for 2 h (the temperature of the reaction was controlled by a sand bath heated at different temperatures). The reaction mixture was monitored by TLC, then filtered to separate the catalyst from the solution. After washing the organic layer with water (3 × 5 mL) and ethyl acetate (25 mL), then making it anhydrous with sodium sulfate, and allowed to crystallize at room temperature. The crystal was filtered and washed with cold ethyl acetate to obtain a pure product. The structure and purity of the product were determined by ¹H, ¹³C NMR, HRMS, and melting point.

5-Amino-9,9-dimethyl-7-oxo-11-phenyl-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (1a). Light yellow powder, Mp. = 242–243 °C, *n*-hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.57. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.36–7.34 (m, 3H), 7.25–7.24 (m, 2H), 7.10–7.08 (m, 2H), 7.07–7.04 (m, 2H), 3.85 (s, 2H), 2.52 (s, 2H), 2.44 (s, 2H), 1.12 (s, 6H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.2, 143.5, 136.9, 134.5, 130.3, 129.6, 129.2, 128.4, 128.1, 127.7, 118.6, 117.1, 109.4, 52.4, 36.9, 35.4, 28.6, 14.3. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 355.1277 found 355.1284.

5-Amino-9,9-dimethyl-7-oxo-11-(*p*-tolyl)-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (2a). White powder, Mp. = 163–164 °C, *n*-hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.59. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.25 (s, 2H), 7.13 (d, *J* = 8.0 Hz, 2H), 7.11–7.09 (m, 2H), 6.93 (d, *J* = 8.0 Hz, 2H), 3.84 (s, 2H), 2.50 (s, 2H), 2.43 (s, 2H), 2.34 (s, 3H), 1.11 (s, 6H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.2, 143.6, 138.4, 134.3, 130.3, 129.8, 129.7, 128.4, 128.0, 127.4, 118.7, 117.0, 109.2, 52.4, 36.9, 35.3, 28.6, 21.1, 14.3. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 369.0446 found 369.0508.

5-Amino-11-(4-methoxyphenyl)-9,9-dimethyl-7-oxo-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (3a). White powder, Mp. = 186–187 °C, *n*-hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.31. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.27 (s, 1H), 7.26 (s, 1H), 7.11–7.09 (m, 2H), 6.97 (d, *J* = 9.0 Hz, 2H), 6.84 (d, *J* = 9.0 Hz, 2H), 3.84 (s, 2H), 3.80 (s, 3H), 2.49 (s, 2H), 2.42 (s, 2H), 1.11 (s, 6H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.2, 159.3, 143.8, 134.7, 130.3, 129.7, 129.6, 128.8, 128.4, 128.0, 118.7, 116.9, 114.4, 109.1, 55.4, 52.4, 36.9, 35.3, 28.6, 14.4. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 385.1915 found 385.1921.

5-Amino-11-(4-bromophenyl)-9,9-dimethyl-7-oxo-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (4a). Light yellow powder, Mp. = 190–191 °C, *n*-hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.50. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.48 (d, *J* = 9.0 Hz, 2H), 7.30–7.28 (m, 2H), 7.10–7.08 (m, 2H), 6.93 (d, *J* =

8.5 Hz, 2H), 3.83 (s, 2H), 2.50 (s, 2H), 2.43 (s, 2H), 1.12 (s, 6H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.1, 143.3, 135.9, 134.4, 132.5, 130.3, 129.2, 128.7, 128.3, 122.4, 118.5, 117.4, 109.8, 52.4, 37.0, 35.4, 28.6, 14.3. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 433.0915 found 433.0920.

5-Amino-11-(3,4-dichlorophenyl)-9,9-dimethyl-7-oxo-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (5a). White powder, Mp. = 227–229 °C, *n*-hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.57. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.41 (d, *J* = 8.5 Hz, 1H), 7.32–7.31 (m, 2H), 7.21 (d, *J* = 2.5 Hz, 1H), 7.11–7.09 (m, 2H), 6.88 (dd, *J* = 2.5 Hz, 8.5 Hz, 1H), 3.83 (s, 2H), 2.52 (s, 2H), 2.44 (s, 2H), 1.13 (s, 6H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.1, 143.2, 136.2, 134.4, 133.4, 133.0, 130.9, 130.3, 129.5, 129.0, 128.8, 128.6, 127.0, 118.3, 117.6, 110.0, 52.3, 36.9, 35.4, 28.6, 14.2. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 423.1030 found 423.1039.

5-Amino-11-(1H-benzo[*d*]imidazol-2-yl)-9,9-dimethyl-7-oxo-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (6a). Yellow liquid, *n*-hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.57. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.61–7.58 (m, 3H), 7.49 (t, *J* = 7.5 Hz, 15.5 Hz, 3H), 7.43–7.39 (m, 2H), 3.99 (s, 2H), 2.81 (s, 2H), 2.43 (s, 2H), 1.19 (s, 6H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.4, 165.6, 151.8, 129.0, 128.8, 126.5, 119.2, 117.1, 107.6, 52.3, 37.4, 35.3, 28.6, 13.5. HRMS(ESI): *m/z* [M-CN]⁺ calcd for 393.2015 found 393.2108.

5-Amino-9,9-dimethyl-7-oxo-11-(*o*-tolyl)-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (7a). White powder, Mp. = 210–213 °C. *n*-Hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.50. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.30–7.28 (m, 1H), 7.24–7.22 (m, 3H), 7.19–7.14 (m, 2H), 7.12–7.10 (m, 2H), 3.92–3.82 (q, *J* = 17.0 Hz, 2H), 2.48–2.38 (m, 3H), 2.19 (d, *J* = 17.0 Hz, 1H), 1.89 (s, 3H), 1.12 (s, 3H), 1.09 (s, 3H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.1, 143.7, 136.0, 135.9, 134.7, 131.1, 130.0, 129.5, 129.3, 128.9, 128.4, 128.2, 126.8, 118.7, 117.0, 109.0, 52.5, 36.5, 35.4, 29.1, 28.0, 17.4, 14.4. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 369.0446 found 369.0508.

5-Amino-7-oxo-11-phenyl-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (1b). White powder, Mp. = 208–210 °C. *n*-Hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.51. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.36–7.32 (m, 3H), 7.25 (d, *J* = 1.5 Hz, 2H), 7.10–7.06 (m, 4H), 3.85 (s, 2H), 2.66 (t, *J* = 6.0 Hz, 12.5 Hz, 2H), 2.56 (t, *J* = 6.0 Hz, 12.5 Hz, 2H), 2.17–2.12 (m, 2H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.9, 144.7, 136.9, 134.3, 130.3, 129.6, 129.2, 128.5, 128.4, 128.1, 127.7, 118.7, 118.3, 109.6, 38.3, 23.6, 23.0, 14.5. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 327.1497 found 327.1495.

5-Amino-7-oxo-11-(3,4-dichlorophenyl)-8,9,10,11-tetrahydro-7H-benzo[*a*]carbazole-6-carbonitrile (5b). White powder, Mp. = 195–198 °C. *n*-Hexane/ethyl acetate = 7/3 (v/v), *R*_f = 0.59. ¹H NMR (500 MHz, CDCl₃): δ (ppm) 7.34–7.32 (m, 3H), 7.11–7.09 (m, 2H), 6.99 (d, *J* = 2.0 Hz, 2H), 3.81 (s, 2H), 2.69 (t, *J* = 6.0 Hz, 12.0 Hz, 2H), 2.56 (t, *J* = 6.0 Hz, 12.0 Hz, 2H), 2.20–2.15 (m, 2H). ¹³C NMR (125 MHz, CDCl₃): δ (ppm) 194.7, 144.3, 138.7, 135.53, 134.2, 130.3, 129.0, 128.9, 128.8, 128.8, 126.4, 118.8, 118.4, 110.4, 38.2, 23.5, 23.0, 14.3. HRMS(ESI): *m/z* [M-CN + 2H]⁺ calcd for 395.1517 found 395.1583.



5-Amino-7-oxo-11-(*o*-tolyl)-8,9,10,11-tetrahydro-7*H*-benzo[*a*]carbazole-6-carbonitrile (7b). White powder, Mp. = 187–192 °C. *n*-Hexane/ethyl acetate = 7/3 (v/v), R_f = 0.49. ^1H NMR (500 MHz, CDCl_3): δ (ppm) 7.25–7.20 (m, 4H), 7.17 (t, J = 7.5 Hz, 1.0 Hz, 2H), 7.12–7.08 (m, 2H), 3.92–3.80 (q, J = 17.0 Hz, 2H), 2.59–2.53 (m, 3H), 2.42–2.35 (m, 1H), 2.17–2.10 (m, 2H). ^{13}C NMR (125 MHz, CDCl_3): δ (ppm) 194.1, 143.6, 136.0, 135.9, 134.7, 131.1, 130.0, 129.5, 129.3, 128.9, 128.4, 128.1, 126.7, 118.7, 117.0, 109.0, 52.5, 36.5, 35.4, 29.1, 28.0. HRMS(ESI): m/z [M-CN + 2H] $^{+}$ calcd for 341.1653 found 341.1655.

Recycling of AC-SO₃H

After the reaction was finished, the product was collected from the reaction mixture using ethyl acetate extraction, and the catalyst was taken out of the solvent. Ethyl acetate was used to wash the recovered catalyst. The model reaction was restarted after it had been dried for two hours at 100 °C. A mixture of (1a) (0.25 mmol) and AC-SO₃H (6 mg) in the presence of DMSO (5.0 mL). The reaction was stirred and refluxed at 240 °C for 2 hours.

Results and discussion

Amorphous carbon-immobilized sulfonic acid characterization (AC-SO₃H)

From the previous research results, AC-SO₃H was constructed using the described method.^{38,39} Rice husk was employed as the starting material, which goes through partial carbonization in an inert atmosphere (known as the amorphous carbon = AC) before being further sulfonated to create AC-SO₃H, which was used as a catalyst. The properties of AC-SO₃H materials were determined by modern technology namely FTIR, P-XRD, TGA, SEM, and EDS. The FTIR spectra of AC-SO₃H and AC are shown in Fig. S1. \dagger The results revealed numerous identifying peaks at 3425 cm^{-1} and 1700 cm^{-1} , corresponding to O–H bonds, and C=O bonds, respectively, which was identified as a signal of the phenolics, carboxylic acids and sulfonic acids site. The signals of the polyaromatic structure, including C=C pairs, were represented by the stretching vibration peak at 1625 cm^{-1} . The AC-SO₃H line exhibits prominent absorption bands of about 1101 cm^{-1} representing the sulfonic acid's S=O stretching vibrations. The results of this study were compared with published studies.⁴⁰

There were three steps in the thermal stabilities of AC shown in Fig. S2, \dagger the first one for the remaining water and finishing at 350 °C. The second stage began at 350 °C and finished at 500 °C due to the decarbonation of amorphous carbon, and the last step began at 550 °C and terminated almost immediately at 800 °C due to the catalyst's evaporation of carbon dioxide gas with a 30% residual. Similar to this, the dehydration of the AC-SO₃H structure was estimated to occur at 500 °C, and the 60% decrease of organic functional groups was estimated to occur at 600 °C. The remaining 35% was kept at a temperature between 600 °C and 800 °C. Thermal gravimetric analysis was used to investigate the thermal stabilities of the waste base material in an N_2 atmosphere.

The SEM image results show that the amorphous carbon material before sulfonation (AC) and after sulfonation (AC-SO₃H) has an agglomerated form (Fig. S3a and b \dagger). Based on EDX analysis of AC and AC-SO₃H materials (Fig. S3d and e \dagger), the structure of AC material contains elements such as C, O, Si, and S with 56.25%, 36.17%, 6.92%, and 0.66% about atomic%, respectively. Besides, AC-SO₃H material also contains elements C (65.56%), O (27.70%), Si (5.19%), and S (1.55%). From the results, the atomic% of the silicon element in AC-SO₃H decreased compared to AC, and the atomic% of the sulfur element in AC-SO₃H increased compared to that of AC-SO₃H with AC. The result showed that the pore structure of AC-SO₃H partially collapsed after sulfonation,⁴¹ and successfully sulfonated AC-SO₃H material. Elemental mapping analysis was also used to confirm the outcomes of the EDX test (Fig. S4 \dagger). This technique also makes it possible to see elemental scattering clearly.

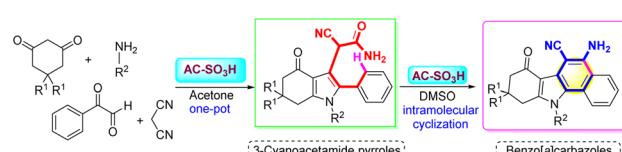
P-XRD analysis was used to characterize the structural characteristics of AC and AC-SO₃H. As can be seen in Fig. S5, \dagger the wide peak C(002), which is associated with the graphitic reflection plane amorphous carbon, can be seen in the diffraction pattern of AC and AC-SO₃H, which is represented by the broad peaks in the XRD pattern, in the 2θ range of 15–35°. The modest and broad C(101) diffraction signal which was observed makes it quite evident that is evidently detected 2θ between 40 and 50° is affected by the graphite structure's a axis.⁴²

The titration method was applied to determine the total acid concentration of AC and AC-SO₃H (Fig. S6 \dagger).

Table S1 \dagger displays the number of surface functional groups determined *via* Boehm titration.⁴³ Using the acidification solution of AC increased the number of total acidic functional groups from an initial value of 0.828 mmol g^{-1} for the AC to 4.606 mmol g^{-1} for the AC-SO₃H. Specifically, greater carboxylic ($-\text{COOH}$) and phenolic ($-\text{OH}$) groups were favored, corresponding with the findings of Mena Aguilar *et al.*⁴⁴ Similar trends were seen between the concentrations of sulfuric acid and $-\text{SO}_3\text{H}$ groups. Inductively Coupled Plasma Mass Spectrometry (ICP-MS) determined that the amount of element S present in the AC-SO₃H catalyst was 2.185 mmol g^{-1} .

Synthesis of benzo[*a*]carbazole frames

AC-SO₃H catalyst activity was evaluated through the synthesis of 3-cyanoacetamide pyrrole from 1,3-diketones, primary amines, phenylglyoxal monohydrate, and malononitrile under acetone, which has been published in our study results.⁴⁵ Continuing to develop from that research result, we investigated the activity of AC-SO₃H as a catalyst for benzo[*a*]carbazole synthesis through



Scheme 2 Synthesis of benzo[*a*]carbazole frames from 3-cyanoacetamide pyrrole scaffolds using AC-SO₃H.



the intramolecular cyclization of 3-cyanoacetamide pyrrole scaffolds to form benzo[*a*]carbazole frames (Scheme 2).

The parameters of the reaction that are investigated when the intramolecular cyclization reaction occurs between 2-cyano-2-(6,6-dimethyl-4-oxo-1,2-diphenyl-4,5,6,7-tetrahydro-1*H*-indol-3-yl)acetamide (**1**) and DMSO to form 5-amino-9,9-dimethyl-7-oxo-11-phenyl-8,9,10,11-tetrahydro-7*H*-benzo[*a*]carbazole-6-carbonitrile (**1a**) include reaction temperature, reaction time, and amount of AC-SO₃H (Table 1). In the presence of 10 mg AC-SO₃H used as a catalyst in DMSO for 4 hours, the reaction was investigated for the influence of reaction temperature with different temperatures extending from 160 °C to 280 °C (entries 1–8, Table 1). The outcomes of the investigation indicate that when the temperature increases, the product (**1a**) also increases, but this process increases slowly and achieves high results at 240 °C with 30%. From the best result, the process of intramolecular cyclization requires high temperature. Next, the reaction time was also surveyed to choose the best time to form the product (**1a**). The reaction time was investigated at 240 °C with different time intervals, including 1 hour, 2 hours, 3 hours, 4 hours, 5 hours, and 6 hours in the presence of 10 mg AC-SO₃H (entries 9–13, Table 1). The results of the trial showed that the best reaction time was two hours. The loading of the AC-SO₃H catalyst added to the reaction mixture was also investigated, consisting of 0, 2, 4, 6, 8, 10, 12, and 14 mg (entries 14–20, Table 1). In cases where the reaction was conducted without a catalyst, the reaction did not occur at all to form the desired product, which is explained by the absence of a source of H⁺ for the oxygen of the carbonyl group, resulting in a reaction that did not

occur. Next, when the amount of catalyst is increased to 2 mg, the reaction was also not released to form a major product (**1a**). However, when AC-SO₃H was used in larger quantities, product formation (**1a**) also increased, and the highest efficiency was achieved when using 6 mg of AC-SO₃H with 73%.

From the results close to the factors affecting the major product formation, the influence of catalyst, solvent, and solvent volume was also evaluated when the intramolecular cyclization reaction occurred with (**1**) (Table 2). In the presence of DMSO as a solvent for the reaction, the activity of the AC-SO₃H catalyst was studied and compared with other catalysts with similar properties to AC-SO₃H. When applying AC as a catalyst for the process, intramolecular cyclization of (**1**) proceeds with poor yield; this is hypothesized to be due to AC only containing -OH and -COOH groups in its structure, which makes its catalytic activity less than AC-SO₃H. IL1-SO₃H was prepared from imidazole,⁴⁶ which was used as the catalyst for the synthesis of (**1a**) with 67%. BAIL gel was synthesized from 1-(4-sulfonic acid)butyl-3-methylimidazolium hydrogen sulfate and tetraethoxysilane,⁴⁷ which was employed to yield of (**1a**) with an isolation yield of 64%. And, SiO₂-H₂SO₄ (ref. 48) was also selected as a catalyst for the reaction because of its similar structure to AC-SO₃H, the reaction yield was as low as 40%. The major product (**1a**) was formed with low yield when using TsOH or H₂SO₄ as catalysts for the intramolecular cyclization reaction of (**1**). At 240 °C, for 2 hours, the reaction did not occur when carried out in the absence of a catalyst or solvent. Under the same reaction conditions, DMSO was replaced by DMF as the solvent for the synthesis reaction (**1a**) with a much lower yield than using DMSO as the reaction solvent, which indicates that DMSO was the best solvent-favorable environment for the reaction due to its ability to be difficult to evaporate and can provide good heat for the reaction. The volume of DMSO solvent was also evaluated for product formation, the volume value investigated included 0, 4, 5, and 6 mL of DMSO solvent. The best results were obtained using 5 mL of DMSO.

Table 1 Effect of the reaction conditions in the synthesis of (**1a**)^a

Entry	Temp. (°C)	Time (h)	Loading of AC-SO ₃ H (mg)	Yields ^b (%)
1	160	4	10	16
2	180	4	10	14
3	200	4	10	22
4	220	4	10	23
5	240	4	10	30
6	250	4	10	33
7	260	4	10	29
8	280	4	10	30
9	240	1	10	34
10	240	2	10	41
11	240	3	10	38
12	240	5	10	33
13	240	6	10	36
14	240	2	0	0
15	240	2	2	0
16	240	2	4	58
17	240	2	6	73
18	240	2	8	74
19	240	2	12	69
20	240	2	14	75

^a Reaction condition: **1** (0.25 mmol, 99.29 mg), AC-SO₃H under DMSO (5 mL). The temperature of the reaction was controlled by a sand bath heated at different temperatures. ^b Yields were isolated yields by crystallization in ethanol (10–15 mL).

Table 2 Evaluate the role of catalysts and solvents for the synthesis of compound (**1a**)^a

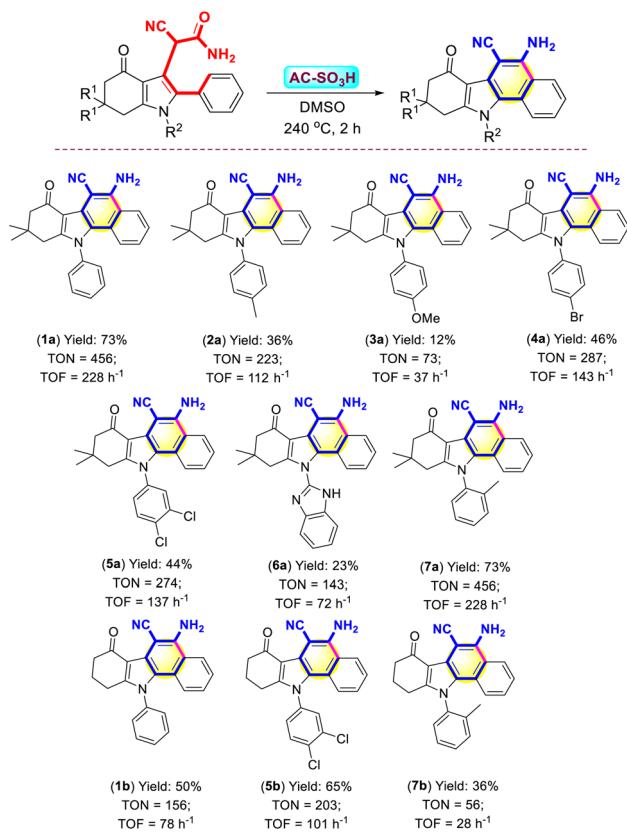
Entry	Catalysts	Solvents	Volume of solvent (mL)	Yields ^b (%)
1	AC	DMSO	5	34
2	AC-SO ₃ H	DMSO	5	73
3	IL1-SO ₃ H ^c	DMSO	5	67
4	BAIL gel ^c	DMSO	5	64
5	SiO ₂ -H ₂ SO ₄	DMSO	5	39
6	TsOH	DMSO	5	25
7	H ₂ SO ₄ ^d	DMSO	5	10
8	None	DMSO	5	0
9	AC-SO ₃ H	None	0	0
10	AC-SO ₃ H	DMF	5	48
11	AC-SO ₃ H	DMSO	4	64
12	AC-SO ₃ H	DMSO	6	58

^a Reaction condition: **1** (0.25 mmol, 99.29 mg), catalyst, and solvents.

^b Yields were isolated yields by crystallization in ethanol (10–15 mL).

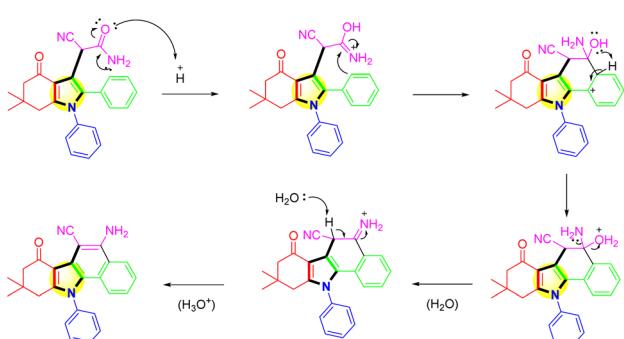
^c IL1-SO₃H = [SiO₂(CH₂)₃Im(CH₂)₄SO₃H]HSO₄; BAIL gel = Brønsted-acidic ionic liquid gel. ^d H₂SO₄ (98.0 wt% in water): 13.0 μL.





Scheme 3 Synthesis of benzo[a]carbazole frames in the presence of AC-SO₃H catalyst. Reaction condition: 3-cyanoacetamide pyrroles (0.25 mmol), AC-SO₃H (6 mg) under DMSO (5 mL) at 240 °C for 2 hours. Yields were isolated yields by crystallization in ethanol (10–15 mL). TON = (moles of product)/(moles of catalyst); TOF = TON/hour (h⁻¹).

The optimal conditions for synthesis, which were established to evaluate the role of the substrate in the synthesis of benzo[a]carbazoles, were 3-cyanoacetamide pyrroles (0.5 mmol), 6 mg of the AC-SO₃H catalyst, the temperature at 240 °C, and DMSO (5 mL) for 2 h. The best result was shown in Scheme 3. By using (1) as the survey object, from the survey results, the isolated yield of forming (1a) reached the highest at 73%. When using (1) carrying electron-donating substituents at the *p*-*p* position, namely –Me and –OMe groups, the major product was low yields 36% and 12%, respectively (Scheme 3, compounds 2a, 3a). However, when



Scheme 4 Mechanism of the intramolecular cyclization of 3-cyanoacetamide pyrrole over the AC-SO₃H catalyst.

the electron-withdrawing substituent was present at the *para* position of (1), the reaction occurred better with over 44% yields, namely –Br, –Cl groups (Scheme 3, compounds 4a, 5a). The pyrrole substrate carries a 1H-benzo[d]imidazol-2-yl substituent, the reaction was difficult due to the resonance of the strong group, reducing the electron density on the carbonyl group of (1), leading to the yield of a major product as low as 23% (Scheme 3, compound 6a). The stereochemistry of the substrate, where the electrons of the –Me group were strongly attracted to the carbonyl group on the active substrate, explained why the electron-donating group at the *ortho* position generated a larger intramolecular cyclization yield than the non-substituent substrate (Scheme 3, compound 7a).

In Scheme 4, the proposed mechanistic pathway for the synthesis of benzo[a]carbazole derivative was shown. For the one-pot reaction of (1), the expected result was attained. 3-Cyanoacetamide pyrrole was hydrolyzed with AC-SO₃H, and the resulting hydrolysis produced an intermediate termed azaniumlidene (A).

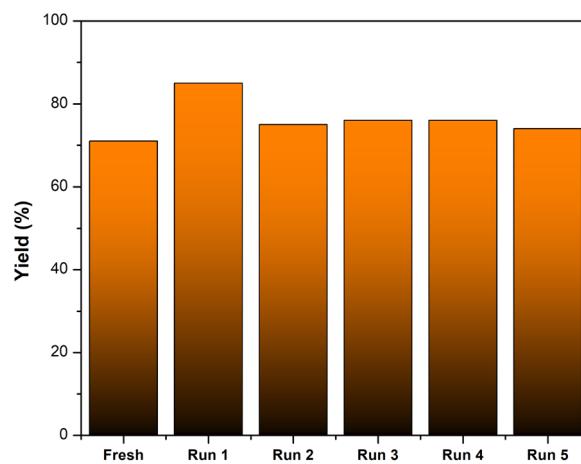


Fig. 1 Recyclability of AC-SO₃H in the synthesis of (1a).

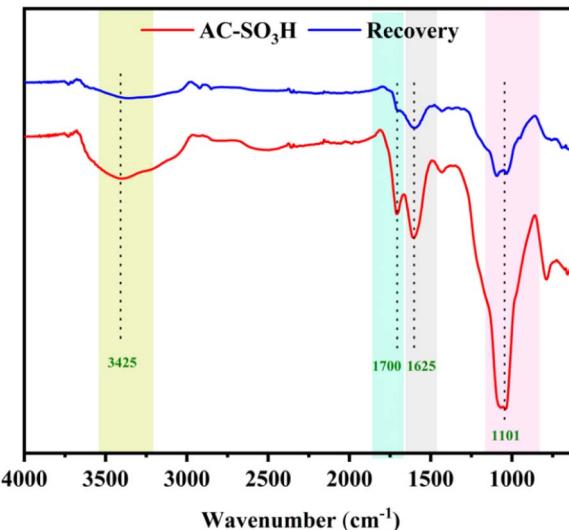


Fig. 2 FT-IR spectrum of AC-SO₃H (fresh) and AC-SO₃H (after 3 times).



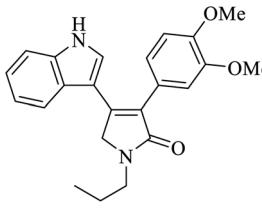
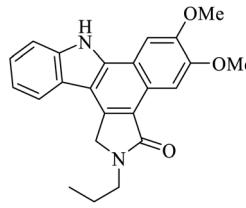
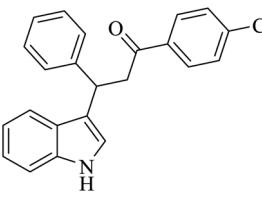
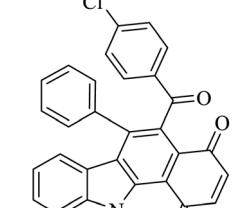
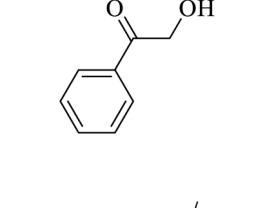
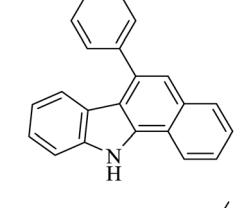
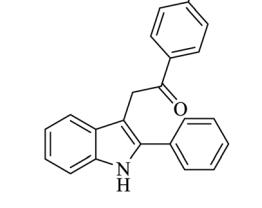
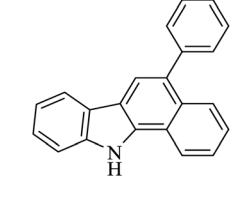
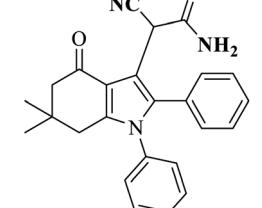
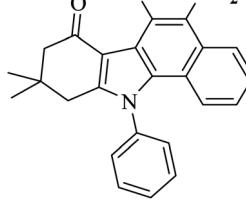
Through the interaction of the azaniumylidene group with the H of ring B, the intermediate (B) was produced by the intramolecular cyclocondensation of intermediate (A). A water molecule was broken down in the reaction to create an intermediate (C). Then protonation to create the major product (**1a**).

The recyclability of AC-SO₃H was evaluated in the intramolecular cyclization of 3-cyanoacetamide pyrrole at 240 °C for 2 hours in the presence of DMSO (Fig. 1). The finished product was separated with ethyl acetate, and the AC-SO₃H was then isolated from the reaction mixture. The recovered catalyst dried for 2 hours at 100 °C and was utilized in the subsequent cycle of the model reaction. Even after five recycling experiments, under optimal conditions, the percentage conversion efficiency

increased from 71% to 74%. The composition of the recovered catalyst was determined by SEM, EDX, and FTIR. As can be seen in Fig. S3,† the agglomeration structure of the AC-SO₃H catalyst after recovery and reuse was still maintained. The collapsible structure of the AC-SO₃H catalyst collapsed after being used three times, and the amount of element sulfur was drastically reduced. From the FTIR spectrum results, the functional group structure in the catalyst was still maintained (Fig. 2).

Table 3 presents a comparison of the benzo[*a*]carbazole framework synthesis method of the study with other studies. The results show that the study's process uses a simpler chemical with a shorter reaction time. However, the reaction requires high temperature.

Table 3 Compare the research method of synthesizing the benzo[*a*]carbazole framework with other research works

Entry	Substrates	Conditions	Product	Yields (%)	Ref.
1		PIFA BF ₃ ·Et ₂ O CH ₂ Cl ₂ −40 °C		43	49
2		Benzoquinone (5 eq.) Toluene/acid acetic (4°:1, v/v) (5 mL) 100 °C, 8 h		76	50
3		TfOH (10% mol) EtNO ₂ , ionic liquid H ₂ O, 95 °C, 4 h		50	51
4		HBr/acid acetic (5% v/v) Argon, 125 °C		69	52
5		AC-SO ₃ H (6 mg) DMSO (5 mL) 240 °C, 2 h		73	Present work



Conclusions

In conclusion, we have developed an eco-friendly and effective catalyst that is made from rice husk. In N_2 atmosphere, carbonization and sulfonation by sulfonic acid were conducted. Immobilizing the sulfonic acid functional group on the carbon framework improves the chemical properties of amorphous carbon compounds. The substance has an aggregated structure according to analytical procedures FT-IR, TGA, XRD, SEM, EDX, and the total acid concentration was found to be 4.606 mmol g^{-1} with elements C (65.56%), O (27.70%), Si (5.19%), and S (1.55%). Through the intramolecular cyclization reaction of 3-cyanoacetamide pyrrole in DMSO at 240 °C for 2 hours, the activity of the AC-SO₃H catalyst was assessed. This study looked at the reaction conditions, and it found that 6 mg of AC-SO₃H increased the intended product yield to 73%. After use, the catalyst was collected and repurposed five times with no structural changes. In addition, compounds containing the benzo[*a*] carbazole framework have many important biological potentials such as glucosidase enzyme inhibition and antibacterial properties.

Author contributions

Hai Truong Nguyen: investigation, methodology, resources, formal analysis, validation, data curation, writing – original draft. Phat Ngoc Nguyen: investigation, methodology, resources, formal analysis, validation. Tan Van Le: investigation, methodology, resources, formal analysis, validation. Trinh Hao Nguyen: investigation, methodology, resources, formal analysis, validation. Linh Dieu Nguyen: investigation, methodology, resources, formal analysis, validation. Phuong Hoang Tran: methodology, resources, formal analysis, validation, data curation, writing – review & editing, supervision.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- 1 H.-Y. Wang, F. Liu, L.-H. Xie, C. Tang, B. Peng, W. Huang and W. Wei, *J. Phys. Chem. C*, 2011, **115**, 6961–6967.
- 2 O. D. Is, F. B. Koyuncu, S. Koyuncu and E. Ozdemir, *Polymer*, 2010, **51**, 1663–1669.
- 3 E. M. Barea, C. Zafer, B. Gultekin, B. Aydin, S. Koyuncu, S. Icli, F. F. Santiago and J. Bisquert, *J. Phys. Chem. C*, 2010, **114**, 19840–19848.
- 4 R. M. Adhikari, D. C. Neckers and B. K. Shah, *J. Org. Chem.*, 2009, **74**, 3341–3349.
- 5 G. Wang, S. Sun and H. Guo, *Eur. J. Med. Chem.*, 2022, **229**, 113999.
- 6 X. Zhu, Y. He, Z. Liu, Z. Zhu, Y. He, J. Qiu, D. Liu, M. Mo, P. Wang, X. Tian and P. Xu, *Appl. Mater. Today*, 2020, **19**, 100559.
- 7 Y. J. Xue, M. Y. Li, X. J. Jin, C. J. Zheng and H. R. Piao, *J. Enzyme Inhib. Med. Chem.*, 2021, **36**, 295–306.
- 8 Y. Chen, N. Cao, H. Lv, K. Zeng, J. Yuan, X. Guo, M. Zhao, P. Tu and Y. Jiang, *Phytochemistry*, 2020, **170**, 112186.
- 9 Y. Tachibana, H. Kikuzaki, N. H. Lajis and N. Nakatani, *J. Agric. Food Chem.*, 2003, **51**, 6461–6467.
- 10 J. A. Scatina, D. R. Hicks, M. Kraml and M. N. Cayen, *Xenobiotica*, 1989, **19**, 991–1002.
- 11 S. Mandal, A. Nayak, M. Kar, S. K. Banerjee, A. Das, S. N. Upadhyay, R. K. Singh, A. Banerji and J. Banerji, *Fitoterapia*, 2010, **81**, 72–74.
- 12 J. H. Burckhalter, V. C. Stephens and L. A. Hall, *J. Am. Pharm. Assoc.*, 1950, **39**, 271–273.
- 13 Y.-P. Liu, J.-M. Guo, Y.-Y. Liu, S. Hu, G. Yan, L. Qiang and Y.-H. Fu, *J. Agric. Food Chem.*, 2019, **67**, 5764–5771.
- 14 R. Birari, S. K. Roy, A. Singh and K. K. Bhutani, *Nat. Prod. Commun.*, 2009, **4**, 1089–1092.
- 15 H.-J. Knölker and K. R. Reddy, *Chem. Rev.*, 2002, **102**, 4303–4428.
- 16 X. Tian, L. Song and A. S. K. Hashmi, *Angew Chem. Int. Ed. Engl.*, 2020, **59**, 12342–12346.
- 17 C. Bosch, P. López-Lledó, J. Bonjoch, B. Bradshaw, P. J. Nieuwland, D. Blanco-Ania and F. P. J. T. Rutjes, *J. Flow Chem.*, 2016, **6**, 240–243.
- 18 S. K. Bur and A. Padwa, *Chem. Rev.*, 2004, **104**, 2401–2432.
- 19 M. F. Martínez-Esperón, D. Rodríguez, L. Castedo and C. Saa, *Tetrahedron*, 2008, **64**, 3674–3686.
- 20 E. Barrera, R. I. Hernández-Benítez, C. A. González-González, C. H. Escalante, A. Fuentes-Benítez, C. González-Romero, E. Becerra-Martínez, F. Delgado and J. Tamariz, *Eur. J. Org. Chem.*, 2022, **2022**, e202200364.
- 21 T. Mandal and J. Dash, *Org. Biomol. Chem.*, 2021, **19**, 9797–9808.
- 22 K. Kirchner, M. J. Calhorda, R. Schmid and L. F. Veiros, *J. Am. Chem. Soc.*, 2003, **125**, 11721–11729.
- 23 A. Banerjee, S. Sahu and M. S. Maji, *Adv. Synth. Catal.*, 2017, **359**, 1860–1866.
- 24 A. F. Saber, S. U. Sharma, J.-T. Lee, A. F. M. El-Mahdy and S.-W. Kuo, *Polymer*, 2022, **254**, 125070.
- 25 S. H. Cho, J. Yoon and S. Chang, *J. Am. Chem. Soc.*, 2011, **133**, 5996–6005.
- 26 W. C. P. Tsang, N. Zheng and S. L. Buchwald, *J. Am. Chem. Soc.*, 2005, **127**, 14560–14561.
- 27 S. Maiti and P. Mal, *Org. Lett.*, 2017, **19**, 2454–2457.
- 28 S. Chen, Y. Li, P. Ni, H. Huang and G.-J. Deng, *Org. Lett.*, 2016, **18**, 5384–5387.
- 29 X. Liu, D. Du, S. Li, X. Wang, C. Xu and M. Wang, *Adv. Synth. Catal.*, 2020, **362**, 5135–5140.
- 30 P. Zhang, B. Li, L. Niu, L. Wang, G. Zhang, X. Jia, G. Zhang, S. Liu, L. Ma and W. Gao, *Adv. Synth. Catal.*, 2020, **362**, 2342–2347.



31 D. Cao, A. Ying, H. Mo, D. Chen, G. Chen, Z. Wang and J. Yang, *J. Org. Chem.*, 2018, **83**, 12568–12574.

32 S. H. Y. S. Abdullah, N. H. M. Hanapi, A. Azid, R. Umar, H. Juahir, H. Khatoon and A. Endut, *Renewable Sustainable Energy Rev.*, 2017, **70**, 1040–1051.

33 L. J. Konwar, P. Mäki-Arvela and J.-P. Mikkola, *Chem. Rev.*, 2019, **119**, 11576–11630.

34 Q. Xie, X. Yang, K. Xu, Z. Chen, B. Sarkar and X. Dou, *Environ. Res.*, 2020, **188**, 109887.

35 S. Ravi and S. Vadukumpally, *J. Environ. Chem. Eng.*, 2016, **4**, 835–856.

36 H. Jirimali, J. Singh, R. Boddula, J. K. Lee and V. Singh, *Materials*, 2022, **15**, 3969.

37 N. Karić, A. S. Maia, A. Teodorović, N. Atanasova, G. Langergraber, G. Crini, A. R. L. Ribeiro and M. Đolić, *Chem. Eng. J. Adv.*, 2022, **9**, 100239.

38 H. T. Nguyen, T. V. Le and P. H. Tran, *J. Environ. Chem. Eng.*, 2021, **9**, 105228.

39 H. T. Nguyen, M.-N. H. Truong, T. V. Le, N. T. Vo, H. D. Nguyen and P. H. Tran, *ACS Omega*, 2022, **7**, 17432–17443.

40 N. Nagasundaram, M. Kokila, P. Sivaguru, R. Santhosh and A. Lalitha, *Adv. Powder Technol.*, 2020, **31**, 1516–1528.

41 S. Na, Z. Minhua, D. Xiuqin and W. Lingtao, *RSC Adv.*, 2019, **9**, 15941–15948.

42 M. Hara, T. Yoshida, A. Takagaki, T. Takata, J. N. Kondo, S. Hayashi and K. Domen, *Angew. Chem., Int. Ed.*, 2004, **43**, 2955–2958.

43 P. V. Rathod and V. H. Jadhav, *ACS Sustainable Chem. Eng.*, 2018, **6**, 5766–5771.

44 K. M. Mena Aguilar, Y. Amano and M. Machida, *J. Environ. Chem. Eng.*, 2016, **4**, 4644–4652.

45 H. T. Nguyen, T. Van Le, P. N. Nguyen, K. H. Nguyen, L. H. T. Nguyen, T. Le Hoang Doan and P. H. Tran, *New J. Chem.*, 2023, **47**, 14733–14745.

46 A. S. Amarasekara and O. S. Owereh, *Catal. Commun.*, 2010, **11**, 1072–1075.

47 P. H. Tran, X.-T. T. Nguyen and D.-K. N. Chau, *Asian J. Org. Chem.*, 2018, **7**, 232–239.

48 D. Kumar, M. Sonawane, B. Pujala, V. K. Jain, S. Bhagat and A. K. Chakraborti, *Green Chem.*, 2013, **15**, 2872–2884.

49 N. J. Truax, F. Banales Mejia, D. O. Kwansare, M. M. Lafferty, M. H. Kean and E. T. Pelkey, *J. Org. Chem.*, 2016, **81**, 6808–6815.

50 C.-W. Kuo, A. Konala, L. Lin, T.-T. Chiang, C.-Y. Huang, T.-H. Yang, V. Kavala and C.-F. Yao, *Chem. Commun.*, 2016, **52**, 7870–7873.

51 M. Li, F. Wu and Y. Gu, *Chin. J. Catal.*, 2019, **40**, 1135–1140.

52 J. Jeon and C.-H. Cheon, *Org. Chem. Front.*, 2019, **6**, 456–467.

