NiO and Co$_3$O$_4$ co-doped g-C$_3$N$_4$ nanocomposites with excellent photoelectrochemical properties under visible light for detection of tetrabromobisphenol-A

Yi Liu,* Jizhou Jiang,* Yanjuan Sun, Shengli Wu, Yuan Cao, Wanyun Gong* and Jing Zou*

Novel NiO/Co$_3$O$_4$-co-doped g-C$_3$N$_4$ nanocomposites with applications in photoelectrochemical sensing were designed and fabricated for the first time in this work. The morphology and microstructure of NiO/Co$_3$O$_4$-co-doped g-C$_3$N$_4$ were comprehensive investigated. The NiO/Co$_3$O$_4$-co-doped g-C$_3$N$_4$ nanocomposites exhibited higher photocurrent and donor density than those of pure g-C$_3$N$_4$, resulting in a higher photoelectrochemical activity for sensitive detection of TBBP-A in real water samples.

Tetrabromobisphenol-A (TBBP-A) is a brominated flame retardant and a persistent organic pollutant. It is widely used in many applications and has caused serious environmental problems and is a great threat to the ecological environment and human health because of its persistence, bioaccumulation and toxicity.\(^1\)–\(^3\) TBBP-A has been mainly detected in air,\(^4\)–\(^5\) water,\(^6\)–\(^7\) sewage sludge,\(^8\)–\(^9\) biological matrices\(^10\)–\(^11\) and soil.\(^12\)–\(^13\) Therefore, it is very important to monitor TBBP-A in organisms and environmental media rapidly and sensitively to assess the human and environmental risks of TBBP-A. Currently, several methods have been used to detect TBBP-A, such as gas chromatography-mass spectrometry (GC-MS), liquid chromatography-mass spectrometry (LC-MS), and tandem mass spectrometry (MS-MS).\(^14\)–\(^16\) Although these methods are effective in detecting TBBP-A accurately, they required large amounts of organic solvent and complicated pretreatment procedures.

In recent years, the photoelectrochemical sensing method, which offers a fast response, low cost and high sensitivity, is an emerging technique for the environmental and biological detection of analytes. Therefore, many high activity photoelectrochemical nanomaterials have been widely investigated. g-C$_3$N$_4$-based nanocomposite, which is a newly developed novel photoelectron nanocomposite, can be used as an electron mediator for effective separation of photo-generated electron-hole pairs at the junction interface.\(^17\) Meanwhile, g-C$_3$N$_4$-based nanocomposites have also exhibited significantly improved optical and electrical activities,\(^18\)–\(^20\) making it one of the most promising materials for the photoelectrochemical sensing method. Therefore, many researchers have made significant efforts to improve the photoelectrochemical activity of g-C$_3$N$_4$ by coupling it with other semiconductors, metals or carbon materials. For example, Cai and co-workers\(^21\) demonstrated a photoelectrochemical methodology for ultrasensitive detection of dopamine based on graphene/g-C$_3$N$_4$/TiO$_2$ nanocomposites, which showed a 0.02 μmol L$^{-1}$ detection limit under the optimized conditions. Our group has also provided an improved method to sensing octylphenol with NiO and Ni co-doped g-C$_3$N$_4$. Our g-C$_3$N$_4$-based nanocomposites modified electrode showed high electrochemical activity to octylphenol under infrared light irradiation.\(^22\) Another useful technique for quantitative assaying of T4 polynucleotide kinase activity based on Au@g-C$_3$N$_4$ nanohybrid was reported by Zhuang.\(^23\) An 2D ternary nano-junction g-C$_3$N$_4$/N-graphene/MoS$_2$ was successfully designed and prepared, exhibiting an enhanced photoelectrochemical performance for simultaneous oxidation of methyl orange and reduction of Cr(VI).\(^24\) Ag modified g-C$_3$N$_4$ film electrodes, which displayed a high photocatalytic activity for the degradation of methylene blue, were successfully prepared by a liquid-based reaction onto the ITO substrates.\(^25\) The successful pairing of photoelectrochemical nanomaterials with a suitable doped catalyst is extremely important because the photoelectrochemical activity may be enhanced by improving the optical absorbance and separating photo-generated carriers at the junction interface. NiO and Co$_3$O$_4$ co-doped catalyst is probably one of the most ideal candidates due to their extraordinary inherent catalytic activity. Up to now, there have been no reports on detection of TBBP-A in real water samples by using photoelectrochemical technique with NiO and Co$_3$O$_4$ semiconductors co-doped g-C$_3$N$_4$ nanocomposites.
In this communication, NiO/Co3O4 co-doped g-C3N4 nano-composites were successfully prepared by a step pyrolysis method, which is simple and low cost-effective. Na film electrodes were fabricated by dropping the mixture solution onto the glassy carbon electrode (GCE). The photoelectrochemical response of the g-C3N4/Na film electrode significantly increased after modifying NiO/Co3O4 co-doped g-C3N4. On the basis of the NiO/Co3O4/g-C3N4 nanocomposites with excellent photo-electrochemical property, a simple and effective photo-electrochemical sensing method with a 0.1 μmol L\(^{-1}\) detection limit for sensitive detection of TBBP-A was developed through differential pulse voltammetry (DPV) under visible light irradiation. The photoelectrochemical sensing method was used for the analysis of TBBP-A in the real samples, and the precision and selectivity of the method were satisfactory.

NiO/Co3O4/g-C3N4 nanocomposites were prepared by pyrolysis method with a small amount of NiCl\(_2\)-6H\(_2\)O and Co(NO\(_3\))\(_3\)-6H\(_2\)O mixtures and 3 g melamine heated at 500 °C for 4 h in a muffle furnace under air atmosphere. The modified electrode of NiO/Co3O4/g-C3N4/GCE was fabricated by dropping the NiO/Co3O4/g-C3N4 nanocomposites suspension on the surface of bare GCE and dried naturally.

The morphology and structure of the NiO/Co3O4/g-C3N4 nanocomposites and g-C3N4 sheets were characterized by TEM. Fig. 1b shows a typically sheet structure of g-C3N4. As shown in Fig. 1a, it can be seen that the NiO and Co3O4 nanoparticles with about 30 nm diameter were dispersed on the ultrathin g-C3N4 sheets. Meanwhile, the high resolution TEM (HR-TEM) image (Fig. 1c) reveals the interface between NiO and Co3O4 nanoparticles. The \(d\)-spacing values of 0.209 and 0.243 nm are clearly observed, which consist of the (200) plane of NiO and (311) plane of Co3O4, respectively. The clear lattice fringes indicate the high crystallinity of the nanocomposites, suggesting the successful preparation of NiO and Co3O4 co-doped g-C3N4 nanocomposites.

The crystalline structure of NiO/Co3O4/g-C3N4 nanocomposites and g-C3N4 sheets were characterized by XRD. In the XRD pattern of g-C3N4, the (002) peak at \(2\theta = 27.4°\) reflected the characteristic interlayer stacking structure, while the (100) diffraction at 13.1° indicated the interplanar structural packing, confirming that the layer-structured g-C3N4 was successfully synthesized. By comparing the XRD patterns of NiO/Co3O4/g-C3N4 nanocomposites and g-C3N4 sheets, it was found that the intensity of the (002) peak at 27.4° significantly decreases and the (100) peak at 13.1° disappears (Fig. 2), indicating a much-lowered long-range order in the atomic arrangements of g-C3N4 in the NiO/Co3O4/g-C3N4 nanocomposites. This result could be attributed to the decreased planar size and structural defects.\(^{26-27}\) We also observed three new peaks at 36.8°, 43.2° and 62.9°, where the former is corresponding to the (311) plane of the cubic phase of Co3O4 (JCPDS no. 43-1003), and the later two peaks are both the characteristic peaks for the cubic phase of NiO (JCPDS 47-1049). These results confirmed that we have successfully prepared the NiO/Co3O4/g-C3N4 hybrid nanocomposites, in accord with the result of the HR-TEM.

The optical absorption property of the materials can considerably determine their photocatalytic activity. The UV-vis diffuse reflectance spectra (DRS) analyses of NiO/Co3O4/g-C3N4, Co3O4/g-C3N4, NiO/g-C3N4 nanocomposites and pure g-C3N4 sheets are shown in Fig. 3a. The main absorption edge of g-C3N4 occurred at 457 nm \((E_g = 2.71\ eV)\), and the NiO/g-C3N4, Co3O4/g-
C3N4, NiO/Co3O4/g-C3N4 nanocomposites are around 470 nm ($E_g = 2.64\,\text{eV}$), 481 nm ($E_g = 2.58\,\text{eV}$), 494 nm ($E_g = 2.51\,\text{eV}$), respectively. This result indicated that NiO-doped and Co3O4-doped both widened the main absorption edge of g-C3N4 and the NiO/Co3O4 co-doped products exhibited the smallest bandgap, improving its conductivity. This result also implied that there was a synergistic effect between NiO and Co3O4 within the NiO/Co3O4/g-C3N4 nanocomposites. Interestingly, there is an extended visible-light absorption over 500–800 nm for all doped nanocomposites, which enables the possible utilization of low-energy visible light, or contributes to the high light-harvesting efficiency of the NiO/Co3O4/g-C3N4 nanocomposites.

To further investigate the photoelectric response behavior of the NiO/Co3O4/g-C3N4 nanocomposites and to understand the role of NiO/Co3O4 in the nanocomposites, the transient photocurrents responses were recorded in 0.1 mol L$^{-1}$ Na$_2$SO$_4$ with repeated light on/off cycles under a 300 W xenon lamp ($\lambda > 420\,\text{nm}$) irradiation (Fig. 3b). NiO/g-C3N4, Co3O4/g-C3N4 and NiO/Co3O4/g-C3N4 nanocomposites all show higher photocurrent densities than that of g-C3N4, suggesting that NiO- or Co3O4-doped can improve the conductivity of g-C3N4. Moreover, the NiO/Co3O4/g-C3N4 nanocomposites displayed a highest photocurrent density among all samples, which indicated that there was a synergistic effect between NiO and Co3O4 within the NiO/Co3O4/g-C3N4 nanocomposites, in concordance with the DRS results (Fig. 3a). These results also indicated that the dramatically increased photocurrent in NiO/Co3O4/g-C3N4 nanocomposites stems from effective charge carrier separation within the nanocomposites via electron transfer.

Additionally, for n-type semiconductor-electrode, the donor density was calculated by the standard formula:

$$\frac{1}{C^2} = \left(\frac{1}{eN_d\varepsilon_0}\right)|V - V_{fb}|$$

where $N_d$ is the donor density (cm$^{-3}$), $\varepsilon_0$ is the permittivity of free space, $V_{fb}$ is the flat-band potential (V), $V$ is the applied potential (V), $\varepsilon$ is the dielectric constant, and $e$ is the electronic charge unit. Meanwhile, Fig. 3c shows the Mott–Schottky plots, typical reversed sigmoid plots of n-type semiconductors, $1/C^2$ versus $E$, for NiO/Co3O4/g-C3N4 nanocomposites and g-C3N4. The $V_{fb}$ of NiO/Co3O4/g-C3N4 nanocomposites and g-C3N4 were estimated to be $-1.53\,\text{V}$ and $-1.12\,\text{V}$ from the Mott–Schottky results. Therefore, the $N_d$ of NiO/Co3O4/g-C3N4 nanocomposites and g-C3N4 could be determined from the slope of the linear region ($2.43 \times 10^{28}\,\text{cm}^{-3}$ and $3.25 \times 10^{27}\,\text{cm}^{-3}$). The $N_d$ of the NiO/Co3O4/g-C3N4 nanocomposites is 7.5-fold that of g-C3N4, indicating the NiO/Co3O4/g-C3N4 nanocomposites possess an improved conductivity. This result also indicated that in comparison with g-C3N4, the electrons could be rapidly transferred at the interface of the NiO/Co3O4/g-C3N4 nanocomposites and electrolyte, which can increase the electrochemical activity of NiO/Co3O4/g-C3N4 nanocomposites.
The electrochemical impedance value is one of the most important electrochemical parameters to understand the electrode materials. Thus, the electrochemical impedance spectroscopy (EIS) was used to characterize the interfacial properties of the nanocomposites in a solution of 0.1 mol L\(^{-1}\) KCl containing 5 mmol L\(^{-1}\) K\(_2\)Fe(CN)\(_6\)/K\(_4\)Fe(CN)\(_6\). The Nyquist plot of the impedance spectra of different modified electrodes are shown in Fig. 3d. Compared with g-C\(_3\)N\(_4\)/GCE and GCE, the smallest semicircles of NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE were observed in the higher frequency region, indicating that NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE had a lowest transfer resistance.

Moreover, we have also explored the photoelectrochemical responses of TBBP-A at the NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE. First, the electrochemical behaviors of TBBP-A at the different modified electrodes were investigated using differential pulse voltammetry (DPV) method in pH = 7.0 phosphate buffered saline (PBS) solution (Fig. 4a). By comparing the oxidation peak current at different modified electrodes, the strongest oxidation peak current of TBBP-A is observed at NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE. The Nyquist plot of the impedance spectra of different modified electrodes are shown in Fig. 3d. Compared with g-C\(_3\)N\(_4\)/GCE and GCE, the smallest semicircles of NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE were observed in the higher frequency region, indicating that NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE had a lowest transfer resistance.

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Hence, pH = 5.8 PBS solution was selected as buffer solution in later experiments. Furthermore, based on the success of the NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\) nanocomposites showing improved photoelectron activity, the electrochemical sensing method for analyzing TBBP-A in water samples was performed under visible-light irradiation as a function of the irradiation time. After two minutes of light irradiation, the oxidation peak current of TBBP-A was 2.5-fold than that of no light irradiation (Fig. 4c). The reason of improved photoelectron activity can be ascribed mainly to the effective separation of electron-hole pairs within NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\) nanocomposites under visible-light irradiation.

In addition, the relationship between peak current and scan rate is powerful tool to better understand the electrochemical performance of TBBP-A at NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE. Thus, cyclic voltammograms (CV) at different scan rates were obtained. As shown in Fig. 5a, the relationship of the anodic (\(I_{pa}\)) peak currents versus the scan rate in the range 20 to 120 mV s\(^{-1}\), whose equation is \(I_{pa} (\mu A) = 0.0219v (V s^{-1}) + 1.4533 (R = 0.995)\), suggesting that a controllable adsorption process occurring at NiO/Co\(_3\)O\(_4\)/g-C\(_3\)N\(_4\)/GCE.

Fig. 5b shows the DPVs of a series of different pH solutions and (c) different lighting times (0.1 mmol L\(^{-1}\) TBBP-A, scan rate: 100 mV s\(^{-1}\)).
suggesting that the NiO/Co3O4/g-C3N4/GCE sensor could be demonstrating good reproducibility. The modifier solution was prepared by mixing 4.0 mg NiO/Co3O4/g-C3N4 dispersed into 4.0 mL of 0.1% NaCl solution. The NiO/Co3O4/g-C3N4 nanocomposites suspension was prepared with 4.0 mg NiO/Co3O4/g-C3N4 nanocomposites dispersed into 4.0 mL of 0.1% NaCl solution. The limit of detection is about 0.1 μmol L$^{-1}$ (S/N = 3). But when the concentration of TBBP-A is more than 500 μmol L$^{-1}$, the electrochemical response becomes low, which could be attributed to the saturation of the electrode surface active sites.

Furthermore, we have investigated the electrochemical sensing performance of the NiO/Co3O4/g-C3N4/GCE in the detection of TBBP-A in real water samples. Two kind of water samples were collected from Campus Lake and Yangtze River with random sampling. As shown in Table 1, the recoveries of TBBP-A in real samples were acceptable recoveries (94.1% - 107.2%).

Meanwhile, the reproducibility of the modified electrode was also investigated by repetitively detecting 0.1 mmol L$^{-1}$ TBBP-A for four times. The relative standard deviation (RSD) was 4.9%, demonstrating good reproducibility. The modified electrode was kept in refrigerator at 4 °C. After four weeks, it retained 92% of the initial peak current response, indicating that NiO/Co3O4/g-C3N4 film have long-term stability. The recoveries were investigated in 0.1 mmol L$^{-1}$ TBBP-A solution in the presence of bisphenol A, o-nitrophenol, m-nitrophenol, K$^+$, CH$_3$COO$^-$ and NO$_3^-$, respectively (Table 2). There was no influence on the peak current, suggesting the high selectivity of the NiO/Co3O4/g-C3N4/GCE for the detection of TBBP-A.

In conclusion, the functional NiO/Co3O4/g-C3N4 nanocomposites were synthesized by a simple method. Compared to g-C3N4, the nanocomposite reflects faster electronic conduction ability and stronger photoelectrochemical activity. An effective and sensitive detection method of TBBP-A was constructed with NiO/Co3O4/g-C3N4/GCE under visible light irradiation. It displayed two linear ranges from 0.3 to 10 μmol L$^{-1}$ and 10 to 500 μmol L$^{-1}$ with a limit of detection of about 0.1 μmol L$^{-1}$.

The NiO/Co3O4/g-C3N4 nanocomposites could be promising for applications in the field of photoelectrochemical sensing.

### Fabrication of the modified electrode

NiO/Co3O4/g-C3N4/GCE was fabricated as follows: NiO/Co3O4/g-C3N4 nanocomposites suspension was prepared with 4.0 mg NiO/Co3O4/g-C3N4 dispersed into 4.0 mL of 0.1% NaCl solution. The solution was then sonicated ca. 10 minutes. Hereafter, bare GCE was mechanically polished to a mirror finish with 0.05 μm alumina slurry. Last, the above nanocomposites suspension of 5.0 μL was dropped on the surface of bare GCE. The modified electrode was dried naturally.

### Characterization

The X-ray powder diffraction (XRD) pattern was determined on an AXS D8 Advance X-ray Diffractometer (Bruker, Germany) equipped with Cu Kα radiation (λ = 1.54 Å) source in the 2θ range from 10 to 80° at a scanning rate of 0.025° s$^{-1}$. The morphologies, structure and particle size of the as-prepared samples were carried out on a JSM-2100 transmission electron microscopy (TEM) (Electronics Co., Japan). The UV-vis diffuse reflectance spectra (DRS) were detected on a Cary Win 50 (Agilent, USA). The photoelectrochemical measurement was performed under a 300 W Xe lamp (Beijing Perfect Co., Ltd.). All electrochemical experiments and photoelectrochemical measurement were performed using a CHI 660E electrochemical workstaton (Shanghai Chenhua Instrument Co., Ltd.) at room temperature. A three-electrode system equipped with a bare GCE or a prepared modified electrode as the working electrode, a Pt wire as the counter electrode and a saturated calomel electrode (SCE) as the reference electrode, was used for all electrochemical measurements.

### Acknowledgements

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### Table 1: Recoveries of TBBP-A in real water samples

<table>
<thead>
<tr>
<th>Sensor</th>
<th>Samples</th>
<th>No</th>
<th>Added</th>
<th>Test</th>
<th>RSD (%)</th>
<th>Recovery (%)</th>
</tr>
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<tbody>
<tr>
<td>NiO/Co3O4/g-C3N4/GCE</td>
<td>Campus Lake water (mol L$^{-1}$)</td>
<td>1</td>
<td>2.50 × 10$^{-6}$</td>
<td>2.60 × 10$^{-6}$</td>
<td>1.94</td>
<td>104.0</td>
</tr>
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<td></td>
<td></td>
<td>2</td>
<td>8.00 × 10$^{-6}$</td>
<td>7.80 × 10$^{-6}$</td>
<td>0.51</td>
<td>97.5</td>
</tr>
<tr>
<td></td>
<td></td>
<td>3</td>
<td>2.00 × 10$^{-6}$</td>
<td>1.88 × 10$^{-5}$</td>
<td>0.26</td>
<td>94.1</td>
</tr>
<tr>
<td>NiO/Co3O4/g-C3N4/GCE</td>
<td>Yangtze River water (mol L$^{-1}$)</td>
<td>1</td>
<td>1.50 × 10$^{-6}$</td>
<td>1.64 × 10$^{-6}$</td>
<td>0.45</td>
<td>109.6</td>
</tr>
<tr>
<td></td>
<td></td>
<td>2</td>
<td>1.80 × 10$^{-5}$</td>
<td>1.72 × 10$^{-5}$</td>
<td>1.49</td>
<td>95.8</td>
</tr>
<tr>
<td></td>
<td></td>
<td>3</td>
<td>2.00 × 10$^{-5}$</td>
<td>1.89 × 10$^{-5}$</td>
<td>1.39</td>
<td>96.8</td>
</tr>
</tbody>
</table>

10 μmol L$^{-1}$ and 10 to 500 μmol L$^{-1}$ (Fig. 5c), which correspond to two linear regression equations: $I_{pa} (\mu A) = 0.0665C$ (μmol L$^{-1}$) + 0.0422 ($R = 0.999$) and $I_{pa} (\mu A) = 0.0062C$ (μmol L$^{-1}$) + 0.7116 ($R = 0.999$), respectively. The limit of detection is about 0.1 μmol L$^{-1}$ ($S/N = 3$). But when the concentration of TBBP-A is more than 500 μmol L$^{-1}$, the electrochemical response becomes low, which could be attributed to the saturation of the electrode surface active sites.

### Table 2: Effects of the interference on the detection of 0.1 mmol L$^{-1}$ TBBP-A

<table>
<thead>
<tr>
<th>Interferences</th>
<th>C (mmol L$^{-1}$)</th>
<th>Signal change (%)</th>
</tr>
</thead>
<tbody>
<tr>
<td>Bisphenol A</td>
<td>0.003</td>
<td>+1.6</td>
</tr>
<tr>
<td>o-Nitrophenol</td>
<td>0.003</td>
<td>+4.9</td>
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<tr>
<td>m-Nitrophenol</td>
<td>0.003</td>
<td>+10.4</td>
</tr>
<tr>
<td>K$^+$</td>
<td>0.3</td>
<td>+5.7</td>
</tr>
<tr>
<td>CH$_3$COO$^-$</td>
<td>0.3</td>
<td>−0.4</td>
</tr>
<tr>
<td>NO$_3^-$</td>
<td>0.3</td>
<td>−7.3</td>
</tr>
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</table>
Notes and references