Introduction

As an emerging class of porous solid materials, metal–organic frameworks (MOFs) that are self-assembled from metal ions or clusters with organic linkers, have attracted considerable attention owing to their great potential in a wide variety of applications. Among the diverse applications of MOFs, gas storage and separation will probably play a key role in future due to the rising energy and environment issues faced by our society. The extremely high surface areas, adjustable pore sizes, as well as intriguing functionalities make them more competitive over other traditional adsorbent materials, such as zeolites, activated carbons, or diatomite. Despite this, the poor stability of MOFs has been recognized to be a major drawback for their practical applications. For this, scientists have proposed some strategies for enhancing the stability of MOFs, how to design and synthesize stable MOFs materials with high capacity of gas adsorption and separation became the focus of research.

A three-dimensional microporous metal–organic framework assembled from a T-shaped pyridyl dicarboxylate ligand, \(4\)’-(pyridin-4-yl)-[1,1’-biphenyl-3,5-dicarboxylic acid (H\(_2\)PBPD), has been successfully synthesized via solvothermal reaction with copper(II) nitrate. Compound 1 was characterized by X-ray single-crystal diffraction, elemental analysis, IR spectroscopy and thermogravimetric analyses. Single-crystal X-ray diffraction analysis indicates that compound 1 is a (3,6)-connected three-dimensional framework with the point symbol \(4^{2}·6\{4\}^{2}·8\{6\}^{2}·10\{3\}\). In the framework of this compound, there exist two large metal–ligand cages. Stability tests show that compound 1 keeps its crystal structure after being soaked in several organic solvents or calcined at 250 °C under air atmosphere. The \(\text{N}_2\) adsorption at 77 K shows that compound 1 is microporous with a BET surface area of 1682 m\(^2\) g\(^{-1}\). Compound 1 possesses high adsorption capacities for \(\text{CO}_2\) (114 cm\(^3\) g\(^{-1}\), 273 K) under ambient pressure. The test of \(\text{CO}_2\) cyclic adsorption and regeneration of 1 shows that it has a stable \(\text{CO}_2\) adsorption capacity and no obvious weight change after doing this ten times.

The structure and properties of MOFs can be rationally designed and improved through several effective approaches such as incorporation of open metal sites, ligand functionalization, framework interpenetration and polar pore walls. Generally, the pores of MOFs are filled with terminal and free solvent molecules, which could be removed by solvent-exchanged and thermal treatment to generate open metal sites. Functionalized organic ligands can be obtained by introducing some groups, such as \(-\text{NH}_3\), \(-\text{OH}\), \(-\text{CH}_3\), and pyridyl groups. And it is worth noting that embedding Lewis basic pyridyl sites into the pore walls of MOFs can drastically impact their gas uptake capacity and selectivity, especially for the \(\text{CO}_2\) capture, which is because there are dipole–quadrupole interactions between the polarizable \(\text{CO}_2\) molecule and the nitrogen site. In addition, the coordination modes of ligands with specific symmetries are a key factor in the structures and properties of the final products.

In recent years, MOFs with carboxylate-containing ligands which contain mixed N-, O-donors have been extensively studied. Among these ligands, the N-donors mainly come from compounds with coordination such as pyridine, imidazole, triazole and so on, and the O-donors come from groups –COOH, –OH or nitrogen oxides. A series of T-shaped pyridyl dicarboxylate ligand such as \(5\)-(pyridin-4-yl)isophthalic acid, \(E\)-5-(pyridin-4-yl-methyleneamino)isophthalic acid, 5-(isonicotinamido)isophthalic acid, 5-[(pyridin-4-ylmethyle)amino]isophthalic acid, 5-(4-
pyridyl)-methoxyl isophthalic acid and 4′-(1H-tetrazol-5-yl)[1,1′-biphenyl]-3,5-dicarboxylic acid (H$_2$TZBPDC), etc. have been used to construct new structures with diverse properties. This type of ligands were selected based on the following considerations: (1) carboxylate groups and pyridine nitrogen donors in the ligand can take part in coordination as bridging ligands to form metal-carboxylate clusters, such as dinuclear, trinuclear, and even higher nuclear clusters; (2) pyridyl donor organic ligands can complement the dicarboxylate ligands in the coordination and formation of high-connected nets; (3) the N-donors from the ligands can coordinate with the metal cations, forming much stronger metal-N bonds, which would enable the high stability of the resultant MOF.

By taking account of the above-mentioned factors, we designed and synthesized a novel T-shaped pyridyl dicarboxylate ligand named 4′-(pyridin-4-yl)[1,1′-biphenyl]-3,5-dicarboxylic acid (H$_2$PBPD) via Suzuki coupling reaction in this work. Based on the H$_2$PBPD ligand, a stable three-dimensional microporous metal-organic framework [Cu(PBPD)(DMF)$_2$(H$_2$O)$_3$] (1) has been successfully synthesized via a solvothermal reaction with copper(n) nitrate. In the framework of this compound, there exist two large metal-ligand cages. We explored the stability and adsorption capacity of compound 1. The stability analyses indicates that compound 1 can keep its crystal structure after being soaked in several organic solvents or calcined at 250 °C under air atmosphere, and even after cycle more than 10 times. The N$_2$ adsorption at 77 K shows compound 1 is microporous with BET surface area of 1682 m$^2$/g. And compound 1 possesses high adsorption capacities for CO$_2$ (114 cm$^3$/g, 273 K) under ambient pressure.

In preparation of this paper, Bai et al. reported a similar MOF showing good CH$_4$ storage capacity under high pressure.90

**Experimental section**

**Materials and physical characterizations**

Except for the H$_2$PBPD ligand, all other chemicals and solvents were obtained from commercial sources and used as received without further purification. Powder X-ray diffraction (PXRD) data were collected on a Rigaku D-Max 2550 diffractometer using Cu-K$_x$ radiation ($\lambda = 0.15418$ nm) and 2$\theta$ ranging from 4 to 40$^\circ$ at room temperature. IR spectra were recorded on a Nicolet Impact 410 FTIR spectrometer as KBr pellets in the 400–4000 cm$^{-1}$ region. Thermogravimetric analyses (TGA) data were recorded on a Perkin-Elmer TGA Q500 thermogravimetric analyzer under an air atmosphere from room temperature to 800 °C (heating rate of 10 °C min$^{-1}$). Elemental microanalyses (C, H and N) were carried out using a Perkin-Elmer 2400 elemental analyzer. The point symbol and topological analysis were conducted using the TOPOS 4.0 program package.

**Crystal structure determination**

The diffraction data of compound 1 were collected on a Bruker AXS SMART APEX-II diffractometer with graphite monochromated Mo K$_x$ radiation ($\lambda = 0.71073$ Å) at 298 K. Data processing was completed with the SAINT processing program. The crystal structures was solved by direct methods, and refined anisotropically on F$^2$ by the full-matrix least-squares technique using the SHELXL-97.11 All non-hydrogen atoms were refined anisotropically. The hydrogen atoms on the aromatic rings were geometrically placed with isotropic thermal parameters 1.2 times that of the attached carbon atom. The diffraction data of compound 1 were collected by using the SQUEEZE routine of PLATON to remove the contribution of disordered molecules.12 Crystallographic data for compound 1 are summarized in Table S1,† while the selected bond lengths and angles are list in Table S2.† More details on the crystallographic studies as well as atom displacement parameters are presents in the CIF.

**Synthesis of 4′-(pyridin-4-yl)[1,1′-biphenyl]-3,5-dicarboxylic acid (H$_2$PBPD)**

The synthetic route of compound 1 is shown in Scheme 1. To a solution of dimethyl 5-iodoisophthalate (320 mg, 1.0 mmol) in DMF (5 mL) were sequentially added 4-bromophenylboronic acid (210 mg, 1.05 mmol), Pd(PhCN)$_2$Cl$_2$ (20 mg, 0.053 mmol) and K$_2$CO$_3$ (400 mg, 2.90 mmol) under N$_2$. After 12 hours stirred at 80 °C under N$_2$, the reaction was quenched with water, extracted with ethyl acetate, washed with brine and dried over MgSO$_4$. Evaporation and flash column chromatography on a silica gel (petroleum ether–ethyl acetate = 10 : 1) gave the coupling product M1 (180 mg, 0.52 mmol) in 52% yield as a pale solid. $^1$H NMR (300 MHz, CDCl$_3$): $\delta$ = 8.66 (t, J = 1.5 Hz, 1H), 8.42 (d, J = 1.8 Hz, 2H), 7.50–7.64 (m, 4H). $^{13}$C NMR (75 MHz, CDCl$_3$): $\delta$ = 166.1, 140.8, 138.0, 132.2, 132.0, 131.4, 129.6, 128.7, 122.7, 52.5. Anal. calcd for C$_{16}$H$_{13}$BrO$_4$: C, 55.04; H, 3.75. Found: C, 55.03; H, 3.51. MS m/z: 351.0 (M + H$^+$, $^{81}$Br), 349.0 (M + H$^+$, $^{79}$Br).

To this product of M1 (710 mg, 2 mmol), dissolved in DMF (10 mL), were added pyridin-4-ylboronic acid (370 mg, 3 mmol), Pd(PPh$_3$)$_2$Cl$_2$ (231 mg, 0.2 mmol) and K$_2$CO$_3$ (552 mg, 4 mmol) under N$_2$. After stirring at 120 °C under N$_2$ for 12 hours, the reaction was quenched with water, recrystallized with ethyl acetate, gave the coupling product dimethyl 4′-(pyridin-4-yl)[1,1′-biphenyl]-3,5-dicarboxylate (M2) (426 mg, 1.22 mmol) in

![Scheme 1 Design and synthesis of ligand H$_2$PBPD.](image-url)
61% yield as a pale solid. 1H NMR (300 MHz, CDCl₃): δ = 8.70 (s, 1H), 8.68 (s, J = 1.5 Hz, 2H), 8.51 (d, J = 1.4 Hz, 2H), 7.82–7.74 (m, 4H), 7.58–7.53 (m, 2H), 3.99 (s, 6H). 13C NMR (75 MHz, CDCl₃): δ = 166.1, 150.4, 147.5, 141.0, 139.8, 138.0, 132.2, 131.4, 129.7, 127.9, 127.7, 121.5, 52.4. Anal. calcd for C₁₉H₁₃NO₄: C, 71.62; H, 4.98, N, 3.36. MS m/z: 348.1 (M + H⁺).

To this product of M₂ (0.96 g, 2.55 mmol), dissolved in THF–H₂O (1:1; 20 mL) was added KOH (1.428 g, 25.5 mmol) and the resulting solution was stirred under reflux for 24 h. The mixture was acidified with aq. HCl, and then filtered. The pure product H₂PBPD was obtained as a pale yellow solid. 1H NMR (300 MHz, DMSO-d₆): δ = 9.00 (d, J = 6.3 Hz, 2H), 8.47–8.55 (m, 5H), 8.21 (d, J = 8.1 Hz, 2H), 8.03 (d, J = 8.4 Hz, 2H). 13C NMR (75 MHz, DMSO-d₆): δ = 166.3, 154.7, 142.0, 141.3, 139.6, 133.9, 132.3, 131.4, 128.9, 128.0, 123.8. Anal. calcd for C₁₉H₁₃NO₄: C, 71.47; H, 4.69%; N, 7.11%. IR spectrum (cm⁻¹): 3428(w), 2928(m), 1668(w), 1632(w), 1599(w), 1585(s), 1521(w), 1493(m), 1470(w), 1372(s), 1348(s), 1297(s), 1265(w), 1229(w), 1202(w), 1139(w), 1104(w), 1048(m), 997(w), 912(w), 819(s), 771(s), 747(s), 649(m), 584(w), 491(m).

Stability test for 1

The as-synthesized sample was soaked in organic solvents such as CH₃OH, C₂H₅OH, CH₂Cl₂, CHCl₃ and CH₃CN for more than two months, then filtered and dried for further characterization of PXRD. In addition, the dried as-synthesized samples were heat at different temperature for 30 min with the heating rate of 5 °C min⁻¹, then the PXRD was also measured to characterize the thermal stability.

Sample activation and adsorption measurements

The solvent-exchange sample was obtained by soaking the as-synthesized sample in methanol for three days with methanol refreshing every two hours, and then replace methanol with dichloromethane in the same method for three days. Before gas adsorption measurement, the sample was future activated using the “outgas” function of the adsorption analyzer to remove the guest molecules from the pores by evacuation under a dynamic vacuum at 85 °C for 10 h. All gases used are of 99.999% purity throughout the adsorption measurements. Low-pressure nitrogen (N₂), carbon dioxide (CO₂) sorption experiments were performed on Micrometrics ASAP 2020 instrument, and methane (CH₄) sorption experiments were performed on Micrometrics ASAP 2420 instrument. The selectivity of CO₂ over CH₄ was calculated by using the ideal adsorption solution theory (IAST). The experiment of CO₂ cyclic adsorption and regeneration of compound 1 were performed on Perkin-Elmer TGA Q500 thermogravimetric analyzer. The sample was pre-treated at 150 °C under N₂ atmosphere for five hours to remove the guest molecules in the pores. The conditions of test was 25 to 80 °C at ambient pressure.

Results and discussion

Crystal structure descriptions

Single-crystal X-ray structural analysis shows that compound 1 possesses 3D neutral framework and crystallizes in the hexagonal space group R̅3. The asymmetric unit of 1 contains one independent Cu atom and one PBPD²⁻ ligand. As shown in Fig. 1a, the two carboxyl groups of PBPD²⁻ ligand display the bridging coordination mode to link two Cu²⁺ ions, and the N atom of pyridine is coordinated with one Cu²⁺ ion. Fig. 1b shows that two Cu²⁺ ions are chelated by four carboxylate groups from four different PBPD²⁻ ligands to form a square paddle-wheel binuclear Cu₂(COO)₄ cluster, and the axial positions of the cluster are occupied by two pyridine-N donors from another two PBPD²⁻ ligands. As shown in Fig. 1c, the six-connected paddle-wheel clusters are linked together through three-connected PBPD²⁻ ligands to form a 3D structure, with 1D channels along the c-axis (8.468 Å). The Cu–O bond lengths varies from 1.954(3) to 1.970(3) Å, the Cu–N bond length is 2.188(3) Å, and the O–Cu–O bond angles are in the range of 88.50(14)–166.82(11)° (Table S2†), which are comparable to those values found in other reported similar compounds. There are two types of metal–ligand cages in the framework of compound 1. The small cage is constructed by six paddle-wheel [Cu₂(CO₂)₄] clusters and six PBPD²⁻ ligands with dimensions of ca. 8.5 × 15.3 Å (Fig. 2a). The large cage consists of twelve paddle-wheel [Cu₂(CO₂)₄] clusters and six PBPD²⁻ ligands with dimensions

![Fig. 1](image-url) (a) Coordination modes of carboxyl groups of PBPD²⁻ ligand. (b) The distorted binuclear [Cu₂(COO)₄] motif. (c) View of the 3D porous network along the c-axis in compound 1. All hydrogen atoms were omitted for clarity.
of ca. 14.9 × 32.1 Å (Fig. 2b). These two cages are arranged alternately to form a cage-stacked 3D framework (Fig. 2c).

Topology analysis is applied in order to better understand the structure of compound 1. If the T-shaped PBPD\(_2\)/C\(_0\) ligand serve as a 3-connected node (Fig. 3a) and the paddle-wheel cluster serve as a 6-connected octahedral node (Fig. 3b), the resulting structure of compound 1 could be simplified as a (3,6)-connected net (Fig. 3c) with the point symbol \{4\^2·6\}\(_2\){4\^4·6\^2·8\^6·10\^3\}, as calculated by Topos 4.0.

**PXRD and thermal analyses of compound 1**

The powder X-ray diffraction (PXRD) shows good correlation between the simulated pattern from the single-crystal data and the as-synthesized pattern of compound 1 (Fig. S2†), indicating that compound 1 is in a high pure phase. In order to identify the thermal stability of compound 1, the thermogravimetric analyses (TGA) have been carried from room temperature to 800 °C (Fig. S4†). The results indicate that compound 1 loses its non-coordinated DMF and H\(_2\)O molecules in the temperature range of 20–334 °C. The weight loss of 34.40% is consistent with calculated one of 34.46%. Then the 3D framework continue to loss weight at higher temperature. The end product of decomposition of compound 1 is CuO with weight percentage 13.62% (calculated: 13.69%).

**Stability of compound 1**

To investigate the chemical stability of compound 1, we soaked the as-synthesized samples in different organic solvents such as CH\(_3\)OH, C\(_2\)H\(_5\)OH, CH\(_3\)CN, CH\(_2\)Cl\(_2\), CHCl\(_3\) for more than two months. And then to investigate the thermal stability, the as-synthesized samples were calcined at different temperature of 150 °C, 200 °C, and 250 °C for 30 min. As shown in Fig. 4, the PXRD pattern of compound 1 treated in different conditions matched the simulated X-ray pattern derived from the single crystal structure, which confirming the structure was not destroyed when soaked in different solvents and calcined in variable temperature. These results indicate that compound 1 possesses high thermal stability. We speculate that the high stability of compound 1 due to the strong energy of the coordination bond based on the Hard-Soft-Acid-Base theory (HSAB), the soft-acid metal cations Cu\(^{2+}\) coordinated with soft-base N donor of H\(_2\)PBPD ligand, forming a much stronger Cu–N bonds. This is the reason why the compound 1 can keep its structure when soaked in different solvents, calcined at high temperature and expose in the air for several months. The HSAB coordination theory maybe an effective strategy for rationally design and synthesize stable MOFs.

**Gas sorption properties**

**N\(_2\) adsorption.** To investigate the permanent porosity of compound 1, the N\(_2\) adsorption experiments were performed at...
77 K. The as-synthesized compound 1 was activated and measured by the adsorption analyze. As shown in Fig. 5, the N₂ adsorption experiments reveals that compound 1 exhibits a typical type-I adsorption isotherm, which is characteristic of microporous materials. The N₂ adsorption shows good reversibility. Compound 1 has a total pore volume of 0.70 cm³ g⁻¹ calculated from the N₂ isotherm data. The Brunauer–Emmett–Teller (S_BET) and Langmuir surface areas (S_Langmuir) are 1682 m² g⁻¹ and 1992 m² g⁻¹, respectively. Under the temperature, the uptake of N₂ was 458 cm³ g⁻¹ at 1.0 atm. The pore size distribution shows that compound 1 possesses three types of pore sizes centered at 0.68 nm, 0.80 nm and 1.00 nm (nonlocal density functional theory method).

**CO₂/CH₄ adsorption.** In addition, low-pressure CO₂ and CH₄ sorption experiments for compound 1 were measured (Fig. 6). The sorption isotherms for CO₂ reveal that compound 1 can absorb CO₂ up to 114 cm³ g⁻¹ (5.08 mmol g⁻¹ or 22.47 wt%) at 273 K and 1.0 atm, and 83 cm³ g⁻¹ (3.70 mmol g⁻¹ or 16.43 wt%) at 298 K and 1.0 atm. The uptake amount of compound 1 is much higher than some of well-known MOFs, such as MOF-5 (6.2 wt%, at 273 K and 1 atm), SUN-4 (20.6 wt%, at 273 K and 1 atm), SUN-21 (18.4 wt%, at 273 K and 1 atm), MOF-602 (5 wt%, at 273 K and 1 atm). However, compound 1 can adsorb 45 cm³ g⁻¹ of CH₄ at 273 K and 1 atm, and 22 cm³ g⁻¹ of CH₄ at 298 K and 1 atm, which is much lower than the uptake of CO₂. To predict the selectivity of CO₂ versus CH₄, we used the ideal adsorbed solution theory (IAST) to calculate the multi-component adsorption behavior from the experimental pure-gas isotherms. The adsorption selectivity for CO₂–CH₄ mixtures as a function of pressure is presented in Fig. S6.† The selectivity of CO₂ over CH₄ reaches a maximum of 6.1 at 298 K and 1.0 bar, indicating that compound 1 exhibits selectivity to CO₂ over CH₄. We speculate that compound 1 can selectively adsorb CO₂ over CH₄ because CO₂ has a large quadrupole moment whereas CH₄ has none.

To better understand the observations and evaluate the extent of CO₂–framework interactions, the Q_st for CO₂ at low coverage are calculated by fitting the gas adsorption isotherms at 273 K and 298 K to the Virial equation. The values are 9.5 kJ mol⁻¹ for CO₂ (Fig. S7†). This low Q_st value shows that the interactions between CO₂ and the framework of compound 1 is weak. The reason may be because the open metal sites of paddlewheel [Cu₄(CO₂)₄] are occupied by the pyridin of PBPD₂⁻ ligand.

Then we explored the CO₂ cyclic adsorption and regeneration of the activated sample under ambient pressure from 25 to 80 °C. As shown in Fig. 7, compound 1 shows a stable CO₂ adsorption capacity and no obvious weight change even after more than 10 cycles. Then the PXRD of the sample after CO₂ adsorption/desorption cycles was measured, as shown in Fig. S2,† the structure of compound 1 was still not destroyed, which also confirmed the stability of the framework.

**Conclusions**

In summary, a metal–organic framework assembled from T-shaped pyridyl dicarboxylate ligand 4′-(pyridin-4-yl)[1,1′-
organic solvents. In addition, the gas adsorption properties of compound 1 are investigated, 

**Notes and references**


**Acknowledgements**

We thank the National Natural Science Foundation of China (Grants: 21471064, 21641004 and 21621001) for support to this work.


11 (a) V. A. Blatov, Struct. Chem., 2012, 23, 955; (b) E. V. Alexandrow, V. A. Blatov, A. V. Kochetkov and D. M. Proserpio, CrystEngComm, 2011, 13, 3947.


