Metallohexacycles containing 4′-aryl-4,2′:6′,4″-terpyridines: conformational preferences and fullerene capture†

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4′-(4-Biphenyl)-4,2′:6′,4″-terpyridine (1) reacts with ZnCl2 or ZnBr2 to produce discrete metallohexacycles instead of the expected one-dimensional coordination polymers. Structural determination of [{ZnCl2(1)}]+ and [{ZnBr2(1)}]+ reveals that the metallomacrocycles adopt a conformation in which the biphenyl domains are in an alternating up/down arrangement (conformer I). The hexamers pack into tubes; within each tube, biphenyl domains of every second hexamer are interdigitated, and these assemblies then interlock to produce a rigid architecture supported by pyridine-phenyl face-to-face contacts. π-Stacking between 4,2′:6′,4″-tpy domains operates between adjacent tubes. Reaction of ZnCl2 or ZnBr2 with 4′-(2′,3′,4′,5′,6′-pentfluorobiphenyl)-4-yl)-4,2′:6′,4″-terpyridine (2) leads to [{ZnCl2(2)}]+ and [{ZnBr2(2)}]+, each crystallizing in two conformations; the centrosymmetric chair-conformer (II) is dominant with respect to the tub-like conformer I. Both conformers pack into tube assemblies, but that consisting of conformer II is less rigid than that of I. Reaction of 4′-(3-naphthalen-1-ylphenyl)-4,2′:6′,4″-terpyridine (3) with ZnCl2 or ZnBr2 leads to [{ZnX(3)}]+ (X = Cl, Br) in conformer I; disordering of the naphthyl substituents is problematic. Assembly of the metallohexacycle in the presence of C60 results in the formation of the host-guest complex [{ZnCl2(3)}]+·C60·3MeOH. The [{ZnCl2(3)}]+ units assemble into a tube-like array that mimics that observed in the parent host. In the host-guest complex, each crystallographically-ordered C60 is trapped between six ordered naphthyl units, three from one hexamer and three from its interdigitated partner, and the C60 six-naphthyl unit sits centrally within a second [{ZnCl2(3)}]+ macrocycle. In contrast to previously described tube-like host-guest assemblies featuring fullerene entrapment, [{ZnCl2(3)}]+·C60 is unusual in having an ordered array of C60 molecules present in every other available cavity, despite the fact that sterically, the ‘empty’ cavity could, in principle, host a C60 guest.

Introduction

A critical strategy in directing the assembly of discrete metallomacrocyclic complexes (metallopolygons) requires complementarity between the coordination preferences of a metal ion and the spatial relationship between the metal-binding domains of the bridging ligand.† Two general approaches are typically adopted. The first uses the coordination geometry of the metal ion to define the internal angle of the metallomacrocycle in combination with rigid-rod ligands, for example, a molecular square predetermined by the 90° angle of square planar platinum(ii).5 The second approach hinges upon the use of a ligand with a fixed internal angle. This strategy has been successfully exploited by the Newkome group to link {M(tpy)}2 units (tpy = 2,2′:6′,2″-terpyridine) into molecular hexagons6 by using ditopic bis(4-substituted tpy) ligands in which the metal ion is effectively a rigid linear metal node. These two design strategies tend to produce metallomacrocycles in which the metal ions are in a roughly planar array. In contrast, judicious choice of ligands can lead to the assembly of helical or grid-like7 macrocycles.4 Regular assemblies can also be obtained with flexible ligands, although a degree of control is lost.

The geometrical flexibility of the zinc(ii) ion (d10) and its compatibility with hard donors (typically N and O) permit zinc(ii) to be applied in a number of ways in multimetallic arrays. Both {Zn2(O2CR)4} and {Zn4(μ-O2CR)4(μ4-O)} building blocks are well-established as nodes with predetermined directionality (linear and octahedral, respectively) in coordination polymers, networks and metal-organic frameworks.
In contrast, combinations of mononuclear zinc(II) nodes (most commonly zinc halides) with bridging ligands containing N- or O-donors lead to a diverse range of coordination polymers and discrete complexes. Reactions of zinc(II) halides with 4,2′:6′,4″-terpyridines tend to lead to coordination polymers (Scheme 1a), but the assembly of a metallohexacycle and polycatenated, triply interlocked metallocapsules has also been observed. The unexpected formation of a molecular metallohexacycle (Scheme 1b) from ZnCl₂ and 4′-(4-ethynylphenyl)-4,2′:6′,4″-terpyridine under crystal growth conditions at ambient temperatures is not readily explained. In contrast to twelve examples of \([\text{ZnX}_2(4′-\text{R′}-4,2′:6′,4″-\text{tpy})]\) \(_n\) \((X = \text{halide or monodentate acetate})\) polymers present in the Cambridge Structural Database (v. 5.34 with November 2012 updates using Conquest v. 1.15), \([\text{ZnCl}_2(4′-(\text{HC}≡\text{CC}_6\text{H}_4)-4,2′:6′,4″-\text{tpy})]_6\) is a unique example of a metallomacrocyclic complex containing a 4,2′:6′,4″-tpy ligand. We now report that this motif is a persistent solid-state product in reactions of ZnCl₂ or ZnBr₂ with three 4′-aryl-4,2′:6′,4″-terpyridines (1–3, Scheme 2). We also describe initial studies of the host–guest chemistry of these metallohexacycles, exemplified by the formation of \([2\text{ZnCl}_2(3)]_6\cdot3\text{MeOH}\).

**Experimental section**

**General**

Electrospray ionisation (ESI) and MALDI TOF mass spectra were measured using Bruker Esquire 3000plus and Bruker microflex instruments, respectively. Solution electronic absorption spectra were recorded using an Agilent 8453 spectrophotometer. \(^1\)H NMR spectra were recorded at room temperature using a Bruker-Avance III-400 spectrometer.

Ligands 1, 2, and 3 were prepared according to literature methods.

\([\text{ZnCl}_2(1)]_6\]

A solution of 1 (19.1 mg, 0.050 mmol) in CHCl₃ (6.0 mL) was placed in a long test tube. MeOH (3.0 mL) was layered on the top of the solution, followed by a solution of ZnCl₂ (6.76 mg, 0.050 mmol) in MeOH (5.0 mL). The test tube was sealed with parafilm and allowed to stand for 3 days at room temperature after which time, colourless crystals had formed. These were isolated by decantation (19 mg, 0.037 mmol, 73%). Found C 61.69, H 4.26, N 7.63%; C₁₆₂H₁₁₄Cl₁₂N₁₈Zn₆·3MeOH requires C 61.42, H 3.94, N 7.81%.

\([\text{ZnBr}_2(1)]_6\]

A solution of 1 (19.1 mg, 0.050 mmol) in CHCl₃ (6.0 mL) was placed in a long test tube and MeOH (3.0 mL) was then layered over the first solution, followed by a solution of ZnBr₂ (11.2 mg, 0.050 mmol) in MeOH (5.0 mL). The tube was sealed with parafilm and allowed to stand for 3 days at room temperature after which time, colourless crystals had formed. These were isolated by decantation (19 mg, 0.037 mmol, 73%). Found C 61.69, H 4.26, N 7.63%; C₁₆₂H₁₁₄Cl₁₂N₁₈Zn₆·3MeOH requires C 61.42, H 3.94, N 7.81%.

\([\text{ZnCl}_2(2)]_6\]

A solution of 2 (23.6 mg, 0.050 mmol) in CHCl₃ (6.0 mL) was placed in a long test tube and MeOH (3.0 mL) was then layered over the first solution, followed by a solution of ZnCl₂ (6.76 mg, 0.050 mmol) in MeOH (5.0 mL). The tube was sealed with parafilm and left to stand at room temperature. Within 3 days, colourless crystals had formed and were isolated by decantation (10.3 mg, 0.0169 mmol, 33.7%). Found C 61.69, H 4.26, N 7.63%; C₁₆₂H₁₁₄Cl₁₂N₁₈Zn₆·2MeOH requires C 52.84, H 3.30, N 6.76%.
Colourless crystals (blocks and spear-like blocks) formed and were isolated by decantation (15.6 mg, 0.0255 mmol, 51.0%).

Found: C 53.62, H 2.31, N 6.87%; C$_2$H$_8$Cl$_2$F$_{30}$N$_{18}$Zn$_6$ requires C 53.02, H 2.81, N 6.87%.

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[[\text{ZnBr}_2(2)]_6]
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A solution of 2 (23.6 mg, 0.050 mmol) in CHCl$_3$ (6.0 mL) was placed in a long tube. MeOH (3.0 mL) was layered on top of the solution, followed by a solution of ZnBr$_2$ (11.2 mg, 0.050 mmol) in MeOH (5.0 mL). The test tube was sealed with parafilm and allowed to stand for 3 days at room temperature. The colourless crystals that formed were isolated by decantation (12.7 mg, 0.0181 mmol, 36.3%). Found C 46.84, H 2.28, N 6.30%; C$_2$H$_{114}$Br$_2$F$_{30}$N$_{18}$Zn$_6$ requires C 46.29, H 2.01, N 6.00%.

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[[\text{ZnBr}_2(3)]_6]
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A solution of 3 (21.8 mg, 0.050 mmol) in CHCl$_3$ (6.0 mL) was placed in a long tube and MeOH (3.0 mL) was layered on top, followed by a solution of ZnBr$_2$ (11.2 mg, 0.050 mmol) in MeOH (5.0 mL). The test tube was sealed with paraffin and left at room temperature for 3 days. The colourless crystals that formed were isolated by decantation (13.6 mg, 0.0206 mmol, 41.2%). Found C 57.38, H 3.81, N 6.21%; C$_{168}$H$_{111}$Cl$_2$F$_{6}$N$_{18}$Zn$_6$ requires C 56.35, H 3.20, N 6.36%.

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2[[\text{ZnCl}_3(3)]_6]_6C_{60}
\]

A solution of 3 (21.8 mg, 0.050 mmol) in a mixture of 1,2-Cl$_2$C$_6$H$_4$ (8.0 mL) and MeOH (2.0 mL) and a solution of C$_{60}$ (6 mg, 0.0008 mmol) in 1,2-Cl$_2$C$_6$H$_4$ (2.0 mL) were placed in a long test tube. A mixture of MeOH (2.5 mL) and 1,2-Cl$_2$C$_6$H$_4$ (2.5 mL) was added as a new layer, followed by a solution of ZnCl$_3$ (6.76 mg, 0.050 mmol) in MeOH (8 mL). After sealing the tube with paraffin, it was left for 2 weeks at room temperature. During this time, purple-red blocks formed in addition to small crystals of C$_{60}$. Satisfactory elemental analysis could not be obtained.

Crystallography

Single crystal data were collected on a Bruker APEX-II diffractometer with data reduction, solution and refinement using the programs APEX2 and SHELXL97 or SHELX-13.27 The ORTEP-type tometer with data reduction, solution and refinement using the Single crystal data were collected on a Bruker APEX-II diffrac-


\[
[[\text{ZnCl}_3(2)]_6]_63\text{CHCl}_3\cdot3\text{MeOH}\cdot6\text{H}_2\text{O (conformer I)}
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C$_{16a}$H$_{111}$Cl$_2$F$_{6}$N$_{18}$O$_{30}$Zn$_{6}$, M = 4232.57, colourless block, trigonal, space group R3, a = b = 37.9984(4), c = 11.3178(11) Å, U = 14152.2(1) Å$^2$, Z = 3, Dx = 1.486 Mg m$^{-3}$, µ(Mo-Kα) = 4.430 mm$^{-1}$, T = 123 K. Total 32865 reflections, 5662 unique, R$_{int}$ = 0.0483. Refinement of 5423 reflections (426 parameters) with I > 2σ(I) converged at final R$_1$ = 0.0613 (R$_1$ all data = 0.0626), wR$_2$ = 0.1810 (wR$_2$ all data = 0.1826), gof = 1.051. CCDC 956347.

\[
[[\text{ZnCl}_2(2)]_6]_6(\text{conformer II)}
\]

After SQUEEZE: C$_{16a}$H$_{111}$Cl$_2$F$_{6}$N$_{18}$O$_{30}$Zn$_{6}$, M = 3670.71, colourless block, monoclinic, space group P2$_1$/n, a = 9.2722(4), b = 35.4234(15), c = 30.2123(12) Å, β = 95.305(2), U = 9880.8(7) Å$^2$, Z = 2, Dx = 1.234 Mg m$^{-3}$, µ(Cu-Kα) = 2.941 mm$^{-1}$, T = 123 K. Total 72063 reflections, 17471 unique, R$_{int}$ = 0.0753. Refinement of 12646 reflections (1182 parameters) with I > 2σ(I) converged at final R$_1$ = 0.1069 (R$_1$ all data = 0.1240), wR$_2$ = 0.3013 (wR$_2$ all data = 0.3169), gof = 1.089. CCDC 956348.

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[[\text{ZnBr}_2(3)]_6]_6\cdot3\text{CHCl}_3\cdot\text{MeOH}\cdot15\text{H}_2\text{O}
\]

C$_{16b}$H$_{159}$Br$_2$Cl$_6$N$_{18}$O$_{15}$Zn$_{6}$, M = 4592.60, colourless block, trigonal, space group R3, a = b = 38.394(4), c = 11.5576(15) Å, U = 141755.3(1) Å$^2$, Z = 3, Dx = 1.419 Mg m$^{-3}$, µ(Cu-Kα) = 4.311 mm$^{-1}$, T = 123 K. Total 29568 reflections, 5754 unique, R$_{int}$ = 0.1096. Refinement of 4031 reflections (343 parameters) with I > 2σ(I) converged at final R$_1$ = 0.0769 (R$_1$ all data = 0.1008), wR$_2$ = 0.1903 (wR$_2$ all data = 0.2024), gof = 1.116. CCDC 956349.

\[
[[\text{ZnBr}_3(3)]_6]_6\cdot3\text{CHCl}_3\cdot15\text{H}_2\text{O}
\]

C$_{16b}$H$_{159}$Br$_2$Cl$_6$N$_{18}$O$_{15}$Zn$_{6}$, M = 4592.60, colourless block, trigonal, space group R3, a = b = 38.3975(2), c = 11.4827(8) Å, U = 15105.8(17) Å$^2$, Z = 3, Dx = 1.505 Mg m$^{-3}$, µ(Mo-Kα) = 3.264 mm$^{-1}$, T = 123 K. Total 125042 reflections, 8934 unique, R$_{int}$ = 0.0580. Refinement of 6705 reflections (585 parameters) with I > 2σ(I) converged at final R$_1$ = 0.0425 (R$_1$ all data = 0.0703), wR$_2$ = 0.1267 (wR$_2$ all data = 0.1717), gof = 1.111. CCDC 956345.

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2[[\text{ZnCl}_3(3)]_6]_6\cdot6\text{MeOH}\cdot16\text{H}_2\text{O}
\]

C$_{43a}$H$_{398}$Cl$_3$N$_{2a}$O$_{2a}$Zn$_{12}$, M = 8062.51, red block, trigonal, space group R3, a = b = 38.4322(16), c = 22.5450(10) Å, U = 28838(2) Å$^3$,
Results and discussion

$[[\text{ZnCl}_2(1)]_6]$ and $[[\text{ZnBr}_2(1)]_6]$

Ligand 1 reacted with ZnCl$_2$ or ZnBr$_2$ under conditions of crystal growth by layering at room temperature. Colourless blocks grew within three days and crystals with only one habit were observed in both reactions. Single crystal X-ray analysis revealed the formation of the molecular metallohexacycles $[[\text{ZnCl}_2(1)]_6]$ and $[[\text{ZnBr}_2(1)]_6]$ as solvates $[[\text{ZnCl}_2(1)]_6]·6\text{CHCl}_3·6\text{MeOH·5H}_2\text{O}$ and $[[\text{ZnBr}_2(1)]_6]·4\text{CHCl}_3·5\text{MeOH·8H}_2\text{O}$. Both compounds crystallize in the trigonal space group $R_3$ with cell parameters that are very similar. The asymmetric unit of each structure contains one molecule of 1 and one ZnX$_2$ unit (X = Cl or Br, Fig. S1 and S2†) and the hexacycle is generated by 3-fold rotoinversion. Fig. 1 shows the structures of the metallocycles $[[\text{ZnCl}_2(1)]_6]$ and $[[\text{ZnBr}_2(1)]_6]$. The two views illustrate the alternating up/down arrangement of the ligands (labelled conformer I), and this contrasts with the three-up/three-down arrangement (corresponding to a chair conformation) observed in $[[\text{ZnCl}_2(1)]_6]·6\text{CHCl}_3·6\text{MeOH·5H}_2\text{O}$.

We return to the question of ring conformation later. In each complex, the Zn atom is tetrahedrally sited and, as is typical of 4,2′:6′,4′-terpyridines, only the outer pyridine rings of ligand 1 are involved in coordination. Bond parameters within the zinc(ii) coordination spheres in solvated $[[\text{ZnCl}_2(1)]_6]$ and $[[\text{ZnBr}_2(1)]_6]$ are given in Table 1. The tpy domain in each complex deviates slightly from planarity (angles between the rings containing N1/N2 and N2/N3 = 5.5 and 14.2° in $[[\text{ZnCl}_2(1)]_6]$, and 4.4 and 12.6° in $[[\text{ZnBr}_2(1)]_6]$, see Fig. S1 and S2† for atom labelling) leading to the bowed backbone seen in Fig. 1a. Twisting of the rings in the biphenyl unit is observed, consistent with relief of H···H repulsions between adjacent rings (angles between the rings containing N2/C16 and C19/C22 are 33.4 and 35.8° in $[[\text{ZnCl}_2(1)]_6]$, and 34.2 and 36.9° in $[[\text{ZnBr}_2(1)]_6]$).

The packing of the hexacycles leads to the formation of a nanotube architecture with the tubes aligned parallel to the crystallographic c-axis (Fig. 2a). The tubes are filled with disordered solvent molecules and crystals are very sensitive to solvent loss, making structure determination difficult for the family of metallohexacycles reported in this work. The organization of molecules within each tube is best described by first considering the interdigitation of pendant phenyl rings of the biphenyl groups of every second metallohexacycle (Fig. 2b). Interlocking of two of the motifs shown in Fig. 2b results in the final nanotubular assembly shown in Fig. 2c. Between the hexacycles coloured red and blue in Fig. 2c, face-to-face π-stacking occurs between the pyridine ring containing atom N1 and the terminal phenyl ring with C22 (symmetry code iii = $1/3 + x, 2/3 + y, 5/3 + z$); in $[[\text{ZnBr}_2(1)]_6]·4\text{CHCl}_3·5\text{MeOH·8H}_2\text{O}$, the angle between the planes is 6.7° and distance between ring centroids is 3.77 Å and analogous parameters are 7.5° and 3.77 Å in the chlorido derivative. As Fig. 3a shows, each hexacycle is involved in six such interactions which contribute significantly...
to the rigid architecture (see later). Adjacent tubes interact through π-stacking of 4,2′:6′:4″tpy domains involving the rings containing N1/N2 and N1iv/N2iv (symmetry code iv = 2/3 − x, 1/3 − y, −2/3 − z) (Fig. 3b). The angle between each pair of stacked pyridine rings is 4.4°, and the inter-centroid distance is 3.68 Å, making this an efficient interaction. Each hexamer is, by symmetry, involved in six such contacts.

The heavily disordered solvent molecules in the two structures have been modelled as partial occupancy H2O, CHCl3 and MeOH, with best models fitting formulations for the compounds of \([\{\text{ZnCl}_2(1)\}_6] \cdot 6\text{CHCl}_3 \cdot 6\text{MeOH} \cdot 5\text{H}_2\text{O}\) and \([\{\text{ZnBr}_2(1)\}_6] \cdot 4\text{CHCl}_3 \cdot 5\text{MeOH} \cdot 8\text{H}_2\text{O}\).

\([\{\text{ZnCl}_2(2)\}_6]\) and \([\{\text{ZnBr}_2(2)\}_6]\)

Despite the fact that the replacement of hydrogen by fluorine in an organic compound can significantly affect solid-state structures,30,31 we have observed that coordination polymers formed between zinc(ii) acetate and ligands 1 and 2 are isostructural.31 We were therefore intrigued to discover whether the same was true for the products of reactions between 1 and 2 with zinc(ii) chloride or bromide. Ligand 2 was combined with ZnCl2 or ZnBr2 under the same room temperature conditions used for reactions with 1. In contrast to the growth of only one type of crystal in reactions with 1, treatment of 2 with either ZnCl2 or ZnBr2 resulted in the formation of colourless blocks and spear-like blocks, the latter being dominant in both reactions.

Structural analysis of the colourless blocks formed from ZnCl2 and 2 confirmed the formation of discrete metallohexacycles in \([\{\text{ZnCl}_2(2)\}_6] \cdot 3\text{CHCl}_3 \cdot 3\text{MeOH} \cdot 6\text{H}_2\text{O}\). The complex crystallizes in the trigonal space group R3 with cell dimensions very similar to those in \([\{\text{ZnCl}_2(1)\}_6] \cdot 6\text{CHCl}_3 \cdot 6\text{MeOH} \cdot 5\text{H}_2\text{O}\). The \([\{\text{ZnCl}_2(2)\}_6]\) hexacycle (Fig. 4a) contains tetrahedral zinc atoms (bond parameters are given in Table 2) and possesses the same conformation as in \([\{\text{ZnCl}_2(1)\}_6]\) with the pentafluorobiphenyl units in an alternating up/down arrangement around the ring (conformer I). Packing of the molecules into tube-like
Assemblies running along the crystallographic c-axis mimics that in \([\{\text{ZnCl}_2(1)\}_6\]·6CHCl_3·6MeOH·5H_2O, with disordered solvent molecules (modelled with partial occupancies) filling the tubes. Packing of \([\{\text{ZnCl}_2(2)\}_6\] hexacycles can be described in the same manner as for \([\{\text{ZnCl}_2(1)\}_6\], but with pentafluorophenyl-pyridine \((\pi_i \cdots \pi_{1i})\) face-to-face \(\pi\)-stacking replacing phenyl-pyridine \(\pi\)-contacts (compare Fig. 4b with Fig. 3a). For the \(\pi_i \cdots \pi_{1i}\)

interaction in \([\{\text{ZnCl}_2(2)\}_6\], the angle between the planes is 7.8° and the distance between ring centroids is 3.71 Å. Packing of tubes involves analogous tpy-tpy face-to-face interactions as detailed for \([\{\text{ZnBr}_2(1)\}_6\] (Fig. 3b); stacked rings lie at 2.2° to one another and the inter-centroid separation is 3.58 Å.

A combination of rapid solvent loss from the block-like crystals formed in the reaction of ZnBr_2 and 2, and heavily disordered solvent, meant that the program SQUEEZE was used to treat the data. Structure determination confirmed the presence of \([\{\text{ZnBr}_2(2)\}_6\] hexacycles. The trigonal space group and cell dimensions were consistent with those determined for all three structures described above, and Fig. 5 shows that the ring adopts conformer I mimicking that in \([\{\text{ZnCl}_2(1)\}_6\], \([\{\text{ZnBr}_2(1)\}_6\] (Fig. 1b) and \([\{\text{ZnCl}_2(2)\}_6\].

The spear-like blocks from the reactions of ZnCl_2 or ZnBr_2 and 2 proved to be a second conformer (conformer II) of the metallohexacycle. The X-ray crystal structure of the chlorido complex \([\{\text{ZnCl}_2(2)\}_6\] confirmed the presence of a chair conformer that replicates that observed in \([\{\text{ZnCl}_2(4)\}_6\] (4 = 4’-(4-ethylnylphenyl)-4,2’;6’,4”-terpyridine). Excessive solvent disorder in the large void space was handled using the program SQUEEZE. \([\{\text{ZnCl}_2(2)\}_6\] crystallizes in the monoclinic space group \(P2_1/n\), with half of the metallomacrocycle in the asymmetric unit; the second half is generated through an inversion centre. Each Zn atom is tetrahedrally coordinated with Zn–N bond distances in the range 2.026(2) to 2.070(2) Å and Zn–Cl bond lengths ranging from 2.2157(8) to 2.2519(10) Å. Fig. 6 shows two views of the structure of \([\{\text{ZnCl}_2(2)\}_6\], and a comparison with Fig. 1 highlights the differences between conformations I and II. One pentafluorobiphenyl unit in conformer II of \([\{\text{ZnCl}_2(2)\}_6\] is disordered and has been modelled over two positions of fractional occupancies 0.79 and 0.21; only one site is shown in Fig. 6. The chair-conformers pack into columns which run parallel to the a-axis (Fig. 7a), with

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**Table 2** Selected bond parameters in \([\{\text{ZnCl}_2(2)\}_6\]·3CHCl_3·3MeOH·6H_2O and \([\{\text{ZnBr}_2(2)\}_6\]. Symmetry codes: \(i = x - y + 2/3, x + 1/3, 1/3 - z\). See Fig. S3 and S4† for atom labelling.

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<th>Bond distance/Å</th>
<th>([{\text{ZnBr}_2(1)}_6]</th>
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**Fig. 4** (a) Structure of the \([\{\text{ZnCl}_2(2)\}_6\] hexamer in \([\{\text{ZnCl}_2(2)\}_6\]·3CHCl_3·3MeOH·6H_2O. (b) Face-to-face stacking of pyridine and pentafluorophenyl rings between adjacent \([\{\text{ZnCl}_2(2)\}_6\] hexamers.

**Fig. 5** Structure of hexamer \([\{\text{ZnBr}_2(2)\}_6\] showing the alternating up/down arrangement of the pentafluorophenyl units (conformer I).
protruding pentafluorophenyl units of one column interdigitated with those of an adjacent column. However, the interdigitation involves pyridine–phenyl $\pi^\mathrm{H} \cdots \pi^\mathrm{H}$ contacts and does not involve the pentafluorophenyl domains. Intermolecular $\pi^\mathrm{F} \cdots \pi^\mathrm{H}$ (pyridine) interactions operate between adjacent $[[\text{ZnCl}_2(2)]_6]$ molecules within a column (Fig. 7a), but only involve one of the three independent $4,2',6',4''$-tpy ligands (that containing N8 and F13, see Fig. S5† for labeling). The angle between the planes through the rings containing N8 and F13 (symmetry code $i = -1 + x, y, z$) is 5.0° and the distance between ring centroids = 3.96 Å. Each hexacycle participates in four such stacking interactions.

Preliminary data only were obtained for the spear-like blocks obtained from reaction of ZnBr$_2$ and 2. Crystals of this habit were repeatedly obtained as the major product in a number of crystallization attempts, but were always of poor quality. Solvent loss was a persistent problem. The preliminary structure determination established the presence of the chair-like conformer of $[[\text{ZnBr}_2(2)]_6]$, thus confirming that $[[\text{ZnBr}_2(2)]_6]$ crystallizes with conformers I and II. Although both conformers pack into tube-like assemblies, the intermolecular interactions between molecules of conformer I, both within a tube and between adjacent tubes, generate a more rigid architecture than those of conformer II. The void spaces (calculated using PLATON$^{29}$) in the lattices of the two conformers of $[[\text{ZnCl}_2(2)]_6]$ are 26.8% for conformer I and 27.6% for conformer II, while for $[[\text{ZnBr}_2(2)]_6]$, the corresponding values are 28.3 and 33.1%.

Interestingly, crystallization by layering a chloroform solution of ZnCl$_2$ with a 1:1 mixture of ligands 1 and 2 in methanol resulted in a compound which crystallizes in the trigonal space group $R3\bar{1}$ with cell dimensions (ESI†) essentially the same as those of $[[\text{ZnCl}_2(1)]_6]$·6CHCl$_3$·6MeOH·5H$_2$O and $[[\text{ZnCl}_2(2)]_6]$·3CHCl$_3$·3MeOH·6H$_2$O. Single crystal X-ray analysis confirmed the assembly of $[[\text{ZnCl}_2(1)]_6]$ in conformation I, with 1 and 2 statistically disordered over one ligand site. We have observed a similar disorder phenomenon in the coordination polymer $[\text{Cu}_2(\mu-\text{OAc})_4(1)]_n\cdot[\text{Cu}_2(\mu-\text{OAc})_4(2)]_n$.$^{24}$

From phenyl to naphthyl: $[[\text{ZnBr}_2(3)]_6]$

In principle, interconversion of conformers I and II of $[[\text{ZnX}_2(1)]_6]$ (L = 1 or 2, X = Cl or Br) may occur through rotation around the Zn–N bonds. The arrangement of the biphenyl or pentafluorobiphenyl units in conformer I appears to
be ideal to host arene guests of appropriate dimensions (e.g. fullerenes), and we argued that a preference for conformer I might be dictated by interactions with suitable guest molecules. Host–guest interactions should be further enhanced by replacing the terminal phenyl or pentafluorophenyl rings in 1 and 2 by more extended aromatic systems, and we chose the naphthyl domain for initial studies (ligand 3, Scheme 2).

The formation of hexamer \( \{[ZnBr_2(3)]_6\} \) was established by single crystal X-ray analysis of colourless blocks of \( \{[ZnBr_2(3)]_6 \cdot 3CHCl_3 \cdot 15H_2O \) that grew from the reaction of ZnBr₂ with 3. Crystallization in the trigonal space group R3 is consistent with the presence of conformer I, and the contents of the asymmetric unit and atom labelling are given in Fig. S6.† Fig. 8 shows the structure of the metallohexacycle. The bond lengths within the tetrahedral coordination sphere of Zn1 are Zn1–N1 = 2.064(2), Zn1–N3 = 2.049(2), Zn1–Br1 = 2.3581(4), Zn1–Br2 = 2.3599(4) Å (symmetry code i = −1/3 + y, 1/3 − x + y, 4/3 − z) with bond angles of N3–Zn1–N1 = 111.28(8) and Br1–Zn1–Br2 = 124.127(18)° and N–Zn1–Br angles in the range 104.02(6) to 108.26(6)°. The naphthalen-1-ylphenyl unit is disordered and has been modelled over 40.8 and 44.3° in the range 104.02(6) to 108.26(6)°. The slight bowing of the tpy backbone (angles between the planes of adjacent pyridine rings = 9.7 and 3.3°) and the twisting of the phenyl ring with respect to pyridine and naphthyl units to which it is bonded (interplane angles = 40.8 and 44.3°) are consistent with the related structures described above. \( \{[ZnBr_2(3)]_6\} \) hexamers stack into tubes along the c-axis in an analogous manner to that detailed in Fig. 2 and the accompanying discussion. Interdigitation of naphthalen-1-ylphenyl units occurs between every second \( \{[ZnBr_2(3)]_6\} \) molecule and adjacent metallohexacycles engage in face-to-face π-stacking of naphthyl and pyridine rings (Fig. 9). The angle between the least squares planes through the pyridine ring containing N3 and the naphthyl unit is 7.3°, and the distances from the centroid of the pyridine ring to those of the rings comprising the naphthyl unit are 3.67 and 4.00 Å. The closest separation of any pair of naphthyl units on one rim of the \( \{[ZnBr_2(3)]_6\} \) hexacycle is ≈11 Å (C29⋯C29’ = 11.1 Å, symmetry code ii = −x + y, 1 − x, z, see Fig. S6† for atom labels) and this compares to the diameter of a C₆₀ molecule of ≈7 Å (10 Å van der Waals diameter). Thus, the cavity is suited to acting as a host for the fullerene.

Preliminary data from structural analysis of crystals grown from reactions of ZnCl₂ and ligand 3 confirmed the assembly of the anticipated hexamer \( \{[ZnCl_3(3)]_6\} \) with conformation I as described for the analogous bromido complex. However, the naphthyl units were heavily disordered, and this was a persistent problem despite modifying the crystallization conditions.

**Fullerene encapsulation:** \( [2\{ZnCl_2(3)\}_{6}C_{60}]6MeOH\cdot16H_2O \)

Porphyrin and calixarene derivatives are popular choices as hosts for C₆₀, with the fullerene typically occupying the bowl-shaped cavity of the host. An example of a dumb-bell shaped metallo-bis(calixarene) (in which two tetrathiacalix[4]arenes are connected through coordination to manganese(ii) centres) breaks this pattern with the C₆₀ guests interacting with the concave outer surface of the dumb-bell in preference to occupying the calixarene cavities. In contrast, both the inner and outer faces of a porphyrin barrel (a tetrameric metalloporphyrin complex) interact with C₆₀, giving rise to a 1:3 host:guest compound in the solid state. Examples of metallomacrocycles hosting C₆₀ appear to be limited. Maverick and coworkers have described C₆₀ encapsulation by a molecular square comprising ditopic β-ketone ligands bound to square planar copper(ii) ions. Several examples of cyclic metallo-bisporphyrins binding fullerene guests have been reported. Of particular relevance to the results reported below is the use of pyridyl-decorated bis(nickellaporphyrin) domains which assemble into one-dimensional tubes in the solid state by virtue of

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**Fig. 8** Structure of the centrosymmetric \( \{[ZnBr_2(3)]_6\} \) molecule in \( \{[ZnBr_2(3)]_6 \cdot 3CHCl_3 \cdot 15H_2O \); the metallohexacycle adopts conformer I.

**Fig. 9** View down the c-axis in \( \{[ZnBr_2(3)]_6 \cdot 3CHCl_3 \cdot 15H_2O \) (solvents omitted) showing face-to-face π-stacking of naphthyl and pyridine domains between adjacent hexamers.
pyridine···pyridine face-to-face π-interactions augmented by CH···Nπpyrrole contacts.\(^4^1\),\(^4^2\)

We considered two approaches for the formation of host–guest complexes using \([\text{ZnX}_2(\text{L})]_6\) metallomacrocycles as hosts. The first strategy of trapping guest molecules within pre-formed hosts relies upon the retention of the metallohexacycles in solution. Unfortunately, we have no unambiguous evidence that the \([\text{ZnX}_2(\text{L})]_6\) complexes remain intact in solution. Attempts to obtain ESI mass spectra were unsuccessful; the MALDI TOF mass spectrum of \([\text{ZnCl}_2(\text{L})]_6\) showed a base (and dominant) peak at \(m/z\: 869.5\) which was assigned to the fragment \([\text{Zn}(\text{L})\text{Cl}]^+\) (calc. \(m/z\: 869.2\)).

The electronic absorption spectrum of a solution made by dissolving crystalline \([\text{ZnCl}_2(\text{L})]_6\) in MeOH (1 \(\times\) 10\(^{-5}\) mol dm\(^{-3}\)) was identical to that of the free ligand,\(^2^3\) suggesting dissociation of the complex. The aromatic region of the \(^1\)H NMR spectrum of a \(\text{CD}_3\text{OD}\) solution of \([\text{ZnCl}_2(\text{L})]_6\) showed a base (and dominant) peak at \(\delta = 8.69\) ppm. A similar result was obtained for a \(\text{CDCl}_3\) solution.

The second strategy for the formation of the host–guest complex involves the assembly of the \([\text{ZnX}_2(\text{L})]_6\) host from \(\text{ZnX}_2\) and \(\text{L}\) in the presence of the guest species. Single crystals suitable for X-ray diffraction were obtained within 2 weeks by carefully layering 1,2-Cl\(_2\)C\(_6\)H\(_4\) crystals suitable for X-ray diffraction were obtained within pre-formed hosts relies upon the retention of the metallohexacycles in solution.

Further, the fullerene–six-naphthyl (green in Fig. 10) assembly lies at the heart of the second (yellow in Fig. 10) \([\text{ZnCl}_2(\text{L})]_6\) hexamer. The \(\text{C}_{60}\) molecule is crystallographically ordered, presumably a consequence of its π-stacking interactions with the naphthyl groups of the host. The closest separations of the centroid of the \(\text{C}_6\)-ring of the fullerene

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**Fig. 10** (a) The two independent \([\text{ZnCl}_2(\text{L})]_6\) hexamers and one \(\text{C}_{60}\) molecule present in \([\text{ZnCl}_2(\text{L})_6\cdot\text{C}_{60}\cdot\text{MeOH}]\cdot\text{H}_2\text{O}\). Only the major occupancy sites of the disordered naphthyl groups (yellow) are shown. (b) Interdigitated naphthalen-1-ylphenyl units of every second hexamer (coloured green) host a \(\text{C}_{60}\) molecule. (c) Interlocking of the crystallographically independent \([\text{ZnCl}_2(\text{L})]_6\) hexamers (coloured green and yellow) completes the structure. Solvent molecules are omitted.
to centroids of the two rings making up the naphthyl group are 3.78 and 4.08 Å. It is noteworthy that only every other set of interdigitated naphthyl units (green in Fig. 10) hosts a fullerene. The cavity between the naphthyl units of the \([{\text{ZnCl}_2(3)}]_6\) hexamers coloured yellow in Fig. 10 is dimensionally similar to that between the hexamers coloured green, but is filled with disordered solvent. The latter have been modelled as partial occupancy \(\text{H}_2\text{O}\) and \(\text{MeOH}\) molecules.

Our attempts to introduce further fullerene into the host (see above) were not successful, begging the question as to why only every other cavity is occupied. The spatial properties of each centro symmetric cavity are essentially the same, and the distance between the middles of empty (yellow in Fig. 10c) and occupied (green) cavities is 11.27 Å. The corresponding distance between the middles of empty (yellow in Fig. 10c) and occupied (green) cavities are essentially the same, and the distance between them is greater than the distance between centroids of the two rings making up the naphthyl group are 3.78 and 4.08 Å. It is noteworthy that only every other set of interdigitated naphthyl units (green in Fig. 10) hosts a fullerene. The cavity between the naphthyl units of the \([{\text{ZnCl}_2(3)}]_6\) hexamers coloured yellow in Fig. 10 is dimensionally similar to that between the hexamers coloured green, but is filled with disordered solvent. The latter have been modelled as partial occupancy \(\text{H}_2\text{O}\) and \(\text{MeOH}\) molecules.

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Conclusions

The formation of metallohexacycles has been shown to be a persistent phenomenon in the reactions of the 4′-aryl substituted 4,2′:6′,4″-terpyridines 1, 2 or 3 with \(\text{ZnCl}_2\) or \(\text{ZnBr}_2\). Whereas \([{\text{ZnCl}_2(1)}]_6\) and \([{\text{ZnBr}_2(1)}]_6\) adopt a conformation (conformer I) in which the biphenyl domains are in an alternating up/down arrangement, both the tub-like conformer I and chair-like conformer II (dominant form) are observed when the pentafluorophenyl domain is present in \([{\text{ZnCl}_2(2)}]_6\) and \([{\text{ZnBr}_2(2)}]_6\). Both conformers stack into tubes in the solid-state, the architecture formed by conformer I being more rigid than that assembled from conformer II. This leads to more robust crystals for conformer I. Hexamers of conformer I pack into tubes. Within each tube, biphenyl domains of every second hexamer are interdigitated, producing motifs which engage in pyridine–phenyl face-to-face contacts forming an interlocked, rigid nanotube. π-Stacking between 4,2′:6′,4″-tpy domains operates between adjacent tubes.

The naphthalen-1-ylphenyl-containing ligand 3 reacts with \(\text{ZnCl}_2\) or \(\text{ZnBr}_2\) to give hexameric \([{\text{ZnCl}_2(3)}]_6\) or \([{\text{ZnBr}_2(3)}]_6\) exhibiting conformer I. If the reaction is carried out in the presence of \(\text{C}_{60}\), crystals of \([2{\text{ZnCl}_2(3)}]_6\cdot\text{C}_{60}\cdot6\text{MeOH·}16\text{H}_2\text{O}\) are isolated. This contains hexacyclic \([{\text{ZnCl}_2(3)}]_6\) molecules in conformer I, replicating the structure of the parent host and its analogue \([{\text{ZnBr}_2(3)}]_6\). The host–guest complex comprises a tube-like structure that mimics that found in \([{\text{ZnCl}_2(3)}]_6\) and \([{\text{ZnBr}_2(3)}]_6\), and in analogous complexes containing \([{\text{ZnX}_2(1)}]_6\) and \([{\text{ZnX}_2(2)}]_6\) in conformer I. Each crystallographically-ordered \(\text{C}_{60}\) is trapped between six ordered naphthyl units, three from one hexamer and three from its interdigitated partner, and the \(\text{C}_{60}\)-six-naphthyl unit sits at the centre of a second \([{\text{ZnCl}_2(3)}]_6\) macrocycle. The structure is highly unusual in having an ordered array of \(\text{C}_{60}\) guests occupying every other available cavity in a tube. All spatial properties of all the six-naphthyl cavities in the lattice are essentially the same, and the distance between them is greater than the separation of \(\text{C}_{60}\) molecules in crystalline \(\text{C}_{60}\) and related structures. Thus, on steric grounds, the ‘empty’ cavity could, in principle, host a fullerene. Thus, we suggest that the observed structure and periodic occupancies of cavities by the fullerene are intimately associated with the assembly process.

Acknowledgements

We thank the Swiss National Science Foundation, the European Research Council (advanced grant 267816 LiLo) and the University of Basel for financial support. We are grateful to Dr. Markus Neuberger for thought-provoking discussions.

Notes and references
