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$CdF(C_6H_4NO_2)(H_2O)$: a UV nonlinear optical material with unprecedented SHG and birefringence *via* π-conjugated rings and a unique "Warren truss structure"†

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We report the design and synthesis of a novel ultraviolet (UV) nonlinear optical (NLO) material, $CdF(C_6H_4NO_2)(H_2O)$, featuring a unique "Warren truss structure". This material exhibits a two-dimensional (2D) layered architecture structure composed of highly polarized $[CdNO_2F_3]$ octahedra and π-conjugated organic rings $(C_6H_4NO_2)^-$. Notably, $CdF(C_6H_4NO_2)(H_2O)$ demonstrates exceptional second-harmonic generation (SHG) response, with an intensity 3.2 times that of KH_2PO_4 (KDP), and a large birefringence of 0.26@546 nm, which is highly unusual for UV fluorides with a bandgap of >4.2 eV. Theoretical calculations and structural analysis reveal that the introduction of $(C_6H_4NO_2)^-$ into CdF_2 induces significant structural distortion and polarization, leading to the formation of a non-centrosymmetric "Warren truss structure". This structure aligns $[CdNO_2F_3]$ octahedra and organic rings in a highly ordered manner, which is crucial for the enhanced SHG and large birefringence. Our findings provide a new strategy for designing high-performance UV NLO materials by leveraging organic–inorganic hybrid structures.

Introduction

The rapid advancement of quantum technology (QT) is revolutionizing the capabilities of communication multiplexing and high-dimensional quantum information processing, with significant implications for spin–orbit angular momentum photonics. In this context, materials with strong second-harmonic generation (SHG) effects and high birefringence are urgently needed to enhance quantum optical techniques, particularly for efficient frequency doubling and precise manipulation of photons carrying orbital angular momentum (OAM).[1–3] However, the development of such materials is hindered by the contrasting microstructural requirements for SHG and birefringence, making it challenging to meet both criteria simultaneously. This is especially true in the ultraviolet (UV) band, where materials with a bandgap of >4.2 eV are required, but few exhibit both strong SHG and large birefringence.

Fluorine, often referred to as the "star element" in nonlinear optical (NLO) materials, is known for its highest electronegativity, which significantly influences crystal structure regulation and optical bandgap enhancement.[4] Its incorporation optimizes birefringence and reduces refractive index dispersion, thereby enabling shorter phase-matching wavelengths.[5] Over the past decade, more than 200 fluorides and their derivatives have been reported, playing a crucial role in optoelectronic applications. However, despite their potential, these materials often fail to combine a large bandgap (>4.2 eV), strong SHG (>3 times KDP), and high birefringence (>0.2).[6–23] Pure metal fluorides, while gaining attention as UV NLO materials, are similarly limited by weak SHG and small birefringence.[24,25] These limitations have significantly restricted the application of fluorides and their derivatives in quantum optical technologies.

Recently, planar conjugated organic rings have garnered significant attention due to their remarkable enhancement of SHG and birefringence. For example, organic groups such as $(H_2C_6N_9)^-$, $(C_7H_4NO_4)^+$, $(C_3N_6H_7)^+$, and $(C_3N_2H_5)^+$ exhibit high optical activity and have been used to construct large birefringent crystals.[26–29] Similarly, groups like $(C_3N_3O_3)^{3-}$, $(C_5H_6NO)^+$, $(HC_3N_3S_3)$, and $(H_2C_6N_7O_3)^-$ have been identified as excellent NLO active units.[30–33] These findings suggest that

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incorporating metal fluorides into organic rings could be an effective strategy to enhance SHG and birefringence. However, this approach remains largely unexplored, particularly for UV-transmitting materials with strong SHG and birefringence.

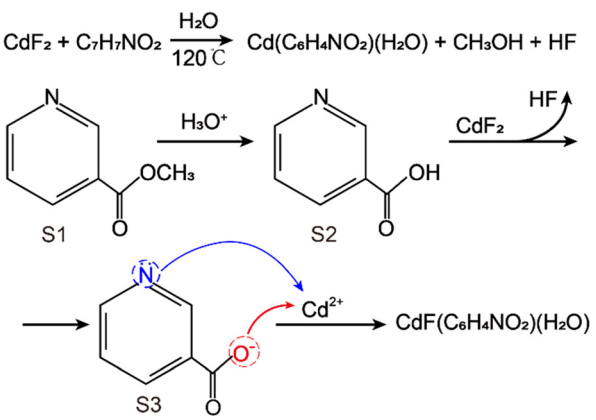
Metal fluoride CdF_2, despite its potential as a UV NLO material, lacks SHG due to its centrosymmetric structure, and its calculated birefringence (0.00002@546.1 nm) is negligible. Here, we report the synthesis of a novel metal–organic complex, $CdF(C_6H_4NO_2)(H_2O)$, by introducing the π-conjugated organic ring $(C_6H_4NO_2)^-$ into CdF_2. This substitution not only breaks the centrosymmetry but also induces significant polarization anisotropy and second-order hyperpolarizability. Compared to the centrosymmetric CdF_2 (SHG = 0, negligible birefringence of 0.00002@546 nm), the resulting compound, $CdF(C_6H_4NO_2)(H_2O)$, exhibits remarkably enhanced SHG (3.2×KDP) and birefringence (0.26@546 nm), while maintaining UV transmittance (band gap = 4.41 eV). This material represents a rare example of UV metal fluoride combining strong SHG and high birefringence. Our comprehensive study, including design, synthesis, structural analysis, properties, and theoretical calculations, reveals how the π-conjugated organic ring $(C_6H_4NO_2)^-$ constructs a "Warren truss structure" in semi-organic metal fluorides. This unique structure aligns highly polarized $[CdNO_2F_3]$ octahedra, leading to the observed strong SHG and large birefringence. Our findings not only advance the understanding of SHG and birefringence enhancement in metal fluorides but also highlight the potential applications of $CdF(C_6H_4NO_2)(H_2O)$ in quantum optical technologies.

Results and discussion

The design idea of $CdF(C_6H_4NO_2)(H_2O)$ is mainly based on the use of $(C_6H_4NO_2)^-$ with a planar π-conjugated organic ring to replace F^- in the simple metal fluoride CdF_2. $CdF(C_6H_4NO_2)(H_2O)$ is obtained using a hydrothermal method as detailed in the ESI.† As shown in Scheme 1, during the reaction process, since CdF_2 is a strong base and weak acid salt, the entire reaction system is acidic. Initially, the raw material methyl nicotinate (S1) hydrolyzes to form nicotinic acid (S2) and methanol. Then, S2 reacts with F^- to produce S3 and the weak acid HF, thereby exposing two coordination active sites of N and O. Subsequently, S3, F^-, and H_2O coordinate with Cd^{2+} to form $CdF(C_6H_4NO_2)(H_2O)$.

Fig. 1a and b illustrate the structural transformation from the precursor CdF_2 to the product $CdF(C_6H_4NO_2)(H_2O)$. The introduction of the organic ring $(C_6H_4NO_2)^-$ disrupts the interconnected $[(Cd_2F_2)^{2+}]_\infty$ chains, forming a unique structure resembling a "Warren truss bridge" within the two-dimensional plane. The rigid organic ring $(C_6H_4NO_2)^-$ acts as the "diagonal" of the bridge, connecting to Cd through coordination atoms at both ends, thereby forming a stable triangular structure that supports the "top/bottom chord" of the bridge, namely the $[(Cd_2F_2)^{2+}]_\infty$ chains, aligning them in a completely parallel and consistent arrangement. The addition of $(C_6H_4NO_2)^-$ transforms the cubic $[CdF_8]$ in CdF_2 into a distorted octahedron $[CdNO_2F_3]$, resulting in structural symmetry breaking from a centrosymmetric to a non-centrosymmetric structure.

$CdF(C_6H_4NO_2)(H_2O)$ crystallizes in the polar space group $P2_1$ with unit cell parameters of a = 9.2128 Å, b = 4.2589 Å, c = 10.740 Å, $\alpha = \gamma = 90°$, β = 114.506(4)°, and V = 383.44(10) Å^3 (detailed information is provided in Table S1†). Its CCDC number is 2419363.† The metal Cd is coordinated with one N and one O from two organic rings $(C_6H_4NO_2)^-$, one water molecule, and three F ions, forming a distorted $[CdNO_2F_3]$ octahedron. The bond lengths of Cd–N, Cd–O, and Cd–F are 2.302 Å, 2.216–2.303 Å, and 2.195–2.311 Å, respectively, which are consistent with previously reported literature (Fig. 2a).[34] As shown in Fig. 2b, the $[CdNO_2F_3]$ octahedra with two orientations are alternately arranged along the b-axis, forming a zigzag one-dimensional chain. These chains are interconnected by the organic rings $(C_6H_4NO_2)^-$ (the distance between two chains is 9.2128 Å, and the angle of the organic ring is 59.2°, as shown in Fig. S1†), creating a unique structure resembling a "Warren truss bridge". They extend infinitely along the ab plane to form a two-dimensional layer. Subsequently, these two-dimensional layers are closely stacked along the c-axis in the same orientation, ultimately forming the unique spatial structure of $CdF(C_6H_4NO_2)(H_2O)$ (Fig. 2c and d).



Scheme 1 Synthesis and mechanism of $CdF(C_6H_4NO_2)(H_2O)$.

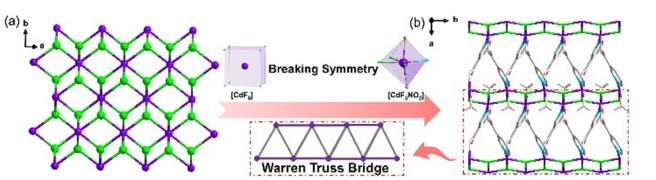


Fig. 1 Transition from a centrally symmetrical CdF_2 structure (a) to a NCS structure of $CdF(C_6H_4NO_2)(H_2O)$ (b). The upper part of the red arrow indicates the evolution of the cube $[CdF_8]$ to the twisted octahedron $[CdNO_2F_3]$. Below the red arrow, a comparison of the $CdF(C_6H_4NO_2)(H_2O)$ structure with the Warren truss bridge is shown.

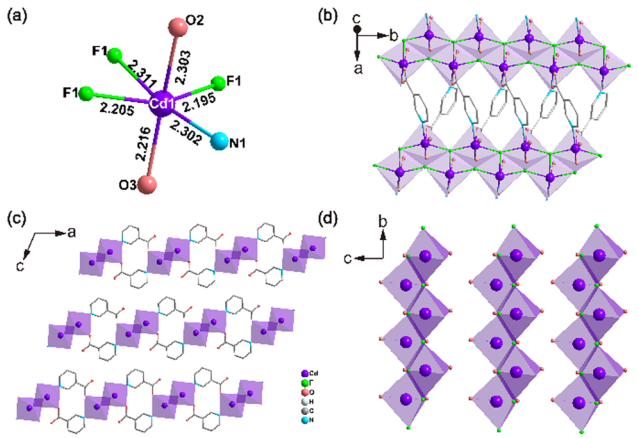


Fig. 2 (a) The coordination environment of Cd; (b) the two-dimensional layered structure of $CdF(C_6H_4NO_2)(H_2O)$ in space; (c) the structure of $CdF(C_6H_4NO_2)(H_2O)$ in the *ac* plane; (d) the structure of $CdF(C_6H_4NO_2)(H_2O)$ in the *bc* plane.

XPS spectral analysis of $CdF(C_6H_4NO_2)(H_2O)$ (Fig. 3a, b and S6, S7†) confirms Cd's oxidation state as Cd^{2+} with peaks at 404.65 eV and 412.62 eV in the 402–414 eV binding energy range and F's oxidation state as −1, consistent with the results of BVS calculations (Table S2†) and single-crystal structure resolution. The band gap of $CdF(C_6H_4NO_2)(H_2O)$ is 4.41 eV, measured using the Kubelka–Munk method (Fig. 3c),[35] exceeding the 4.2 eV threshold for UV NLO materials and surpassing reported values for Na_2CeF_6 (3.89 eV),[36] KBi_4F_{13} (4.24 eV),[37] and $K_2SbF_2Cl_3$ (4.01 eV),[38] indicating its UV application potential. IR absorption peaks align with single-crystal test results (Fig. S3†), and thermogravimetric tests show stability up to 150 °C (Fig. S4†).

The Kurtz–Perry method was used to measure[39] the SHG response of $CdF(C_6H_4NO_2)(H_2O)$ under 1064 nm laser irradiation. Its SHG intensity increases with particle size, plateaus at 280–450 μm, and shows phase matching behavior. Within this size range, its SHG intensity is about 3.2×KDP (Fig. 3d and e). Generally, metal fluorides have weak NLO effects due to fluorine's weak deformability. For instance, the SHG responses of KNa_2ZrF_7 (0.35×KDP),[4] $CsNaTaF_7$ (0.20×KDP),[40] $BaMgF_4$ (0.085×KDP),[41] $BaZnF_4$ (0.16×KDP),[42] Na_2SbF_5 (0.17×KDP),[43] and Na_2CeF_6 (2.1×KDP)[36] are all less than 1×KDP. Although Na_2CeF_6 has a rare 2.1×KDP SHG response, its 3.89 eV optical band gap limits its UV band application. Fig. 3f shows the comparison of the optical band gaps and SHG effects of recent metal fluorides (Table S8†), showing that $CdF(C_6H_4NO_2)(H_2O)$'s SHG effect is the strongest among those of recent UV NLO metal fluorides with a band gap of >4.2 eV. Further dipole moment analysis reveals a calculated dipole moment of 6.14 D for $CdF(C_6H_4NO_2)(H_2O)$, aligning with its SHG intensity.

The second harmonic generation (SHG) effect is generally diminished in fluorides with short absorption edges. The origin of the nonlinear optical (NLO) efficiency is predominantly contingent on the asymmetry and arrangement of the polyhedra within the crystal structure. In the compound $CdF(C_6H_4NO_2)(H_2O)$, the severely distorted octahedra $[CdNO_2F_3]$ constitute a one-dimensional sawtooth chain. Subsequently, these chains are interconnected by the organic ring

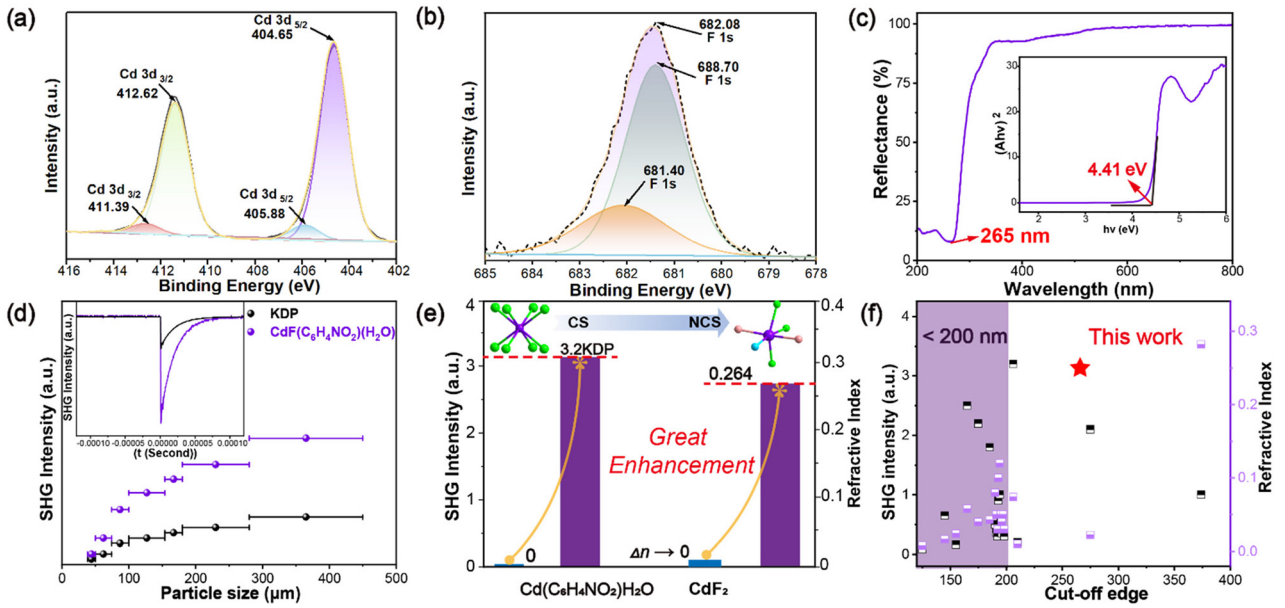


Fig. 3 (a) XPS spectrum of Cd-3d in $CdF(C_6H_4NO_2)(H_2O)$. (b) XPS spectrum of F-1s in $CdF(C_6H_4NO_2)(H_2O)$. (c) UV spectrum of $CdF(C_6H_4NO_2)(H_2O)$. (d) The particle size of $CdF(C_6H_4NO_2)(H_2O)$ and KDP as a function of the SHG response. The inset is a SHG intensity signal plot of $CdF(C_6H_4NO_2)(H_2O)$ with KDP in the particle size range of 280–450 μm. (e) Comparison of SHG response and birefringence performance of $CdF(C_6H_4NO_2)(H_2O)$ and CdF_2. (f) Scatter plots of bandgap, birefringence and SHG intensities of $CdF(C_6H_4NO_2)(H_2O)$ *versus* various fluorides. (Black on the left indicates SHG, purple on the right represents birefringence, and black at the bottom denotes the cut-off edge.)

$(C_6H_4NO_2)^-$, culminating in a distinctive "Warren truss bridge" structure. This unique configuration facilitates the optimal alignment of the $[CdNO_2F_3]$ octahedra and $(C_6H_4NO_2)^-$ within the crystal lattice, which is the pivotal factor enabling the large SHG and high birefringence of $CdF(C_6H_4NO_2)(H_2O)$. These structural analyses offer invaluable insights into the design of NLO materials with enhanced SHG effects, potentially guiding future research endeavors in this domain.

The birefringence of $CdF(C_6H_4NO_2)(H_2O)$ single crystals, measured using a polarizing microscope (Fig. 4a), was found to be 0.253@546.1 nm for a crystal thickness of 7 μm (Fig. 4b–d). This value exceeds those of commercial birefringent crystals such as MgF_2 (0.012@589.3 nm),[44] α-BaB_2O_4 (0.122@532 nm),[45] and $CaCO_3$ (0.172@589 nm)[46] and represents the upper limit of UV pure metal fluorides and semi-organometallic fluorides (Table S8†), except for $(H_2DpA)_2SiF_6$(0.282).[47] $(H_2DpA)_2SiF_6$ has only 1×KDP and a band gap of 2.84 eV, significantly lower than the required 4.2 eV and 3×KDP. The calculated linear optical properties (Fig. 4e) reveal strong anisotropy with refractive indices n_z = 1.764, n_x = 1.745, and n_y = 1.511, yielding birefringence Δn = 0.264@546 nm, agreeing well with the measured value. Additionally, the birefringence of CdF_2 at 546 nm was calculated to be 0.00002 (Fig. 4f), which is nearly 10 000 times lower than that of $CdF(C_6H_4NO_2)(H_2O)$.

First-principles density-functional theory calculations (Fig. S8–S11†) reveal that $CdF(C_6H_4NO_2)(H_2O)$ and CdF_2 have indirect band gaps of 3.507 and 3.743 eV, respectively. The density of states diagram (Fig. S11†) shows that the valence band top is dominated by O-2p orbitals with minor F-2p contributions, while the conduction band bottom is influenced by C-2p, N-2p, and O-2p orbitals. This indicates that the band gap of $CdF(C_6H_4NO_2)(H_2O)$ is primarily determined by the $(C_6H_4NO_2)^-$ unit, with minimal contribution from Cd–F interactions (Fig. 5b).

Quantum chemistry-based frontier orbital calculations on $CdF(C_6H_4NO_2)(H_2O)$ primitives reveal the contributions of $[CdNO_2F_3]$ and $(C_6H_4NO_2)^-$ units. In $CdF(C_6H_4NO_2)(H_2O)$, the HOMO is dominated by F-2p and Cd-4d orbitals, while the LUMO is primarily influenced by the $(C_6H_4NO_2)^-$ unit and Cd and F orbitals within the $[CdNO_2F_3]$ octahedron (Fig. S12 and S13†). These d–p hybridization events facilitate electron movement, under the photoelectric field, enhancing the second-harmonic generation (SHG) effect.

Under Kleinman symmetry constraints,[48] $CdF(C_6H_4NO_2)(H_2O)$ exhibits four independent nonzero SHG coefficients: d_{14} = −0.241 pm V^{-1}, d_{16} = 1.114 pm V^{-1}, d_{22} = 0.747 pm V^{-1}, and d_{23} = −0.439 pm V^{-1}, with d_{16} being the largest and consistent with experimental results. Further investigation through SHG-weighted density maps (Fig. 5a, b and S14, S15†) shows that the occupied states in the virtual electron (VE) and virtual hole

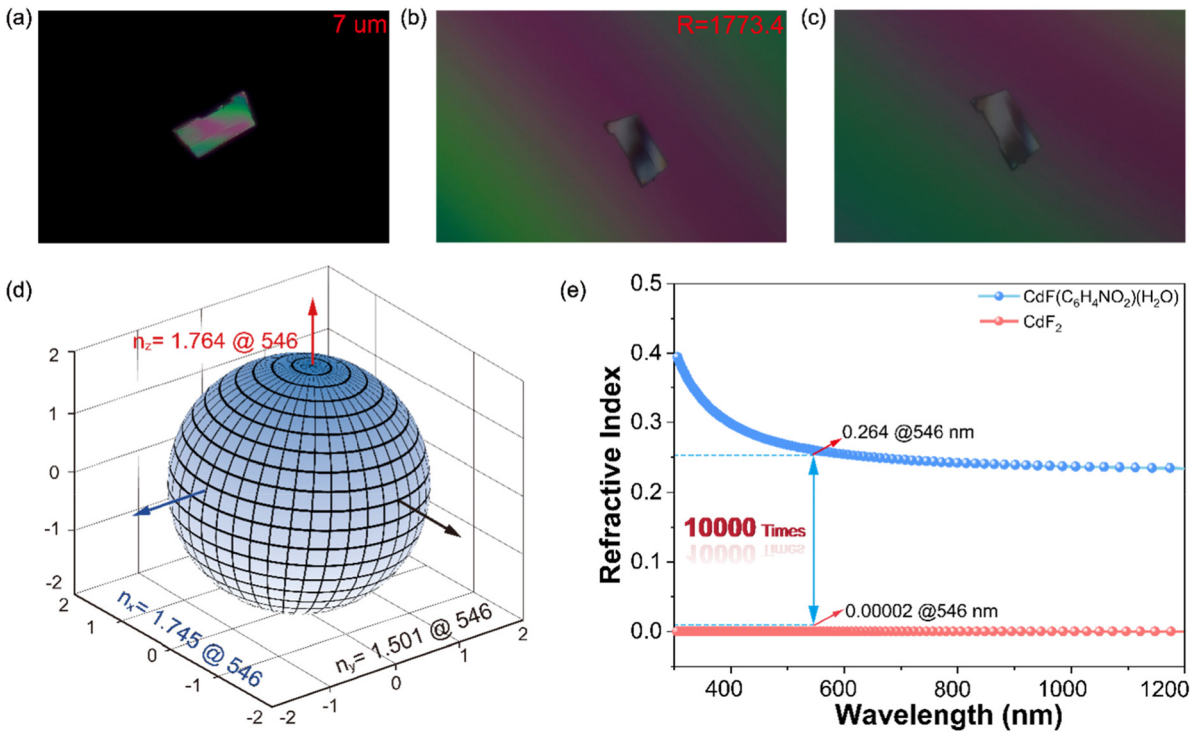


Fig. 4 (a) Thickness of selected wafers measured using the birefringence of $CdF(C_6H_4NO_2)(H_2O)$. (b and c) $CdF(C_6H_4NO_2)(H_2O)$ extinguished under cross-polarized light. (d) Theoretically calculated refractive index of $CdF(C_6H_4NO_2)(H_2O)$. (e) Comparison of birefringence@546 nm between CdF_2 and $CdF(C_6H_4NO_2)(H_2O)$.

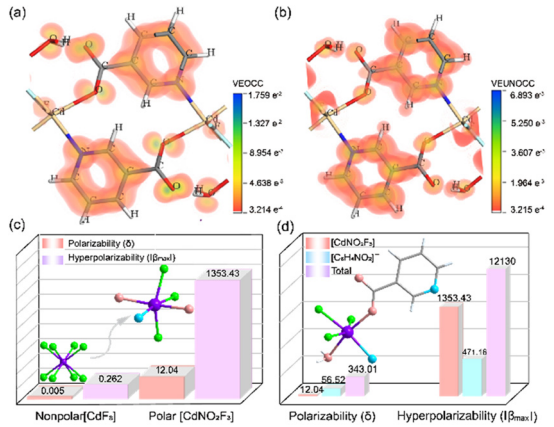


Fig. 5 SHG-weighted densities of the occupied (a) and unoccupied (b) states of $CdF(C_6H_4NO_2)(H_2O)$ in the virtual electron process. (c) Hyperpolarizability and polarization anisotropy calculations for the non-polar octahedron $[CdF_8]$ and the polar octahedron $[CdNO_2F_3]$. (d) Theoretical calculation of anisotropy and hyperpolarizability contributions of $CdF(C_6H_4NO_2)(H_2O)$, $[CdNO_2F_3]$ and $(C_6H_4NO_2)^-$.

(VH) processes are mainly from C-2p, N-2p, O-2p, and F-2p orbitals, while the unoccupied states are primarily from C-2p, N-2p, and Cd-4d orbitals. These findings confirm that the SHG density arises from the synergistic contributions of the $(C_6H_4NO_2)^-$ unit, F, and Cd atoms.

It is well known that anisotropy and hyperpolarizability are key parameters affecting second harmonic generation (SHG) and birefringence, respectively. We calculated these properties for $[CdF_8]$, $[CdNO_2F_3]$, $CdF(C_6H_4NO_2)(H_2O)$, and $(C_6H_4NO_2)$ using the LanL2DZ basis set in Gaussian[49] (Fig. 5c and d). For CdF_2, $[CdF_8]$ shows near-zero hyperpolarizability and anisotropy, consistent with its poor birefringence and lack of SHG. In contrast, $[CdNO_2F_3]$ in $CdF(C_6H_4NO_2)(H_2O)$ exhibits hyperpolarizability and anisotropy values over 2000 and 5000 times higher than those of $[CdF_8]$, highlighting its crucial role in enhancing optical properties (Fig. 5c). Fig. 5d shows that $(C_6H_4NO_2)^-$ significantly contributes to birefringence, while the highly polarized $[CdNO_2F_3]$ units drive strong SHG. Taken together, it is confirmed that the unique "Warren truss structure" formed by the π-conjugated organic ring $(C_6H_4NO_2)^-$ leads to highly polarized and well-aligned $[CdNO_2F_3]$ octahedra and $(C_6H_4NO_2)^-$, which is the key factor leading to the strong SHG and birefringence of $CdF(C_6H_4NO_2)(H_2O)$.

Conclusion

In summary, we have successfully synthesized $CdF(C_6H_4NO_2)(H_2O)$, a novel UV NLO material with a unique "Warren truss structure". Compared to CdF_2, which exhibits zero SHG effect and extremely low birefringence (0.00002@546 nm), $CdF(C_6H_4NO_2)(H_2O)$ demonstrates remarkable enhancements in both properties, achieving a large SHG response (3.2×KDP) and high birefringence (0.26@546 nm). These striking improvements are attributed to the introduction of the π-conjugated organic ring $(C_6H_4NO_2)^-$, which leads to the formation of a highly polarized and well aligned $[CdNO_2F_3]$ octahedra. Notably, this is the first UV fluoride material to simultaneously exhibit a large band gap (>4.2 eV), strong SHG effect (>3×KDP), and high birefringence (>0.2), making it a promising candidate for quantum optical technologies. Our study provides valuable insights into the design of non-centrosymmetric optoelectronic materials by constructing novel structures that activate high polarization and enhance optical properties.

Data availability

The data supporting this article have been included as part of the ESI.†

Conflicts of interest

The authors declare no competing financial interest.

Acknowledgements

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