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Exploration of antimony(III) oxyhalides via single-site substitution in quest of large birefringence†

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Birefringent crystals play a crucial role in the laser polarization of modern laser technologies. As a marvelous branch of optical materials, oxyhalides are attracting extensive interest for their suitable structures and diverse properties. Metal cations with lone pairs have proven advantageous for enhancing birefringence and extending the range of transmission. In this study, we comprehensively investigated the antimony(III) oxyhalide system, Sb–O–X. Specifically, we systematically examined the impact of single-site substitution within a series of compounds, including SbOCl, Sb₂OCl₄, Sb₃O₄F, Sb₃O₄I, Sb₃O₄Cl, Sb₈O₁₁Cl₂, and Sb₈O₁₁Br₂. The substitution of halogens led to significant alterations in the crystal structures, ranging from 0D isolated units to 2D layers, which are favorable for generating birefringences greater than 0.1. These findings underscore the potential of antimony oxyhalides for achieving the balance between birefringence and bandgap, and affirm the viability of single-site substitution as an effective strategy for discovering birefringent materials.

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Introduction

The polarization of light, as an extraordinary natural phenomenon, has yielded numerous discoveries and a wide array of applications. Crystals with great birefringence have played a pivotal role in both research and engineering domains for manipulating the polarization of light, encompassing applications in polarimetry, optical fibers, and the laser industry. The discovery of the new birefringent materials system has attracted considerable academic and commercial attention. As nonlinear optical (NLO) materials continue to advance rapidly, there is a pressing need for birefringent materials due to their essential technological attributes in the manipulation of light

polarization.^{1–7} Several birefringent materials, including MgF₂,⁸ α-BaB₂O₄,^{9,10} CaCO₃,¹¹ TiO₂,¹² and YVO₄ crystals,¹³ have been harnessed for applications spanning from the deep-ultraviolet (DUV) to the near-infrared (NIR) regions. With the increased demand for birefringent materials, materials and systems with large birefringence still deserve to be discovered. Consequently, scientists continually explore birefringent functional units and structures to achieve great birefringence.^{14,15}

It is widely recognized that birefringence is intricately linked to the anisotropic polarizability of material structure. To harness this characteristic, researchers have often selected π-conjugated planar groups, such as (NO₃)[–], (CO₃)^{2–}, (BO₃)^{3–}, (B₃O₆)^{3–}, (C₃N₆)^{6–}, (C₄O₄)^{2–}, and (C₃N₃O₃)^{3–}, as ideal motifs for designing birefringent materials due to their notably large anisotropic linear polarizability.^{16–37} In contrast to their purely oxidized tetrahedral counterparts, non-π-conjugated heterotetrahedra, such as [BO_xF_{4–x}](x = 0–4), [SO₃F], [PO_xF_{4–x}](x = 2, 3), [PO_xS_{4–x}](x = 0–4), and [SiO_xN_{4–x}](x = 0–4), have been reported to exhibit an increased birefringence.^{38–47} These findings suggest that the exploration of birefringent materials can be broadened to a much broader horizon by finding new birefringent units.

In recent years, there has been a significant surge of interest in oxyhalides due to their applications in energy, environmental processes, ion conductivity, and photocatalytic activities.^{48–50} Oxyhalides are also emerging as promising candidates for the IR region, primarily because of their

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exceptional resistance to laser-induced damage (LD).^{51–57} Compared to pure metal oxides and halides, oxyhalides exhibit a more favorable balance between bandgap and birefringence. The incorporation of halogens leads to structural distortions and increases polarizability anisotropy. Furthermore, the introduction of cations with lone pairs, such as Sb^{3+} , Bi^{3+} , Pb^{2+} , and Sn^{2+} , proves advantageous for extending the IR cutoff edge and achieving greater birefringence when compared to the spherical-shaped alkali and alkali earth metal ions.^{58–69} A recent example of this phenomenon is the about 60 times enhancement of the birefringence in $\alpha\text{-NaSb}_3\text{P}_2\text{O}_{10}$ over $\text{K}_3\text{SbP}_2\text{O}_9$. This remarkable enhancement is attributed to the presence of the Sb^{3+} cations with stereochemically active lone pair in $\alpha\text{-NaSb}_3\text{P}_2\text{O}_{10}$, while the Sb^{5+} cations in $\text{K}_3\text{SbP}_2\text{O}_9$ carry no lone pair.⁶⁴ These findings underline the significant contribution of Sb^{3+} cations with lone pairs to birefringence. As a result, oxyhalides with variable antimony(III)-centered polyhedra would open the door to a multitude of fascinating structural possibilities.

While the incorporation of $[\text{Sb-O-X}]$ ($\text{X} = \text{halide}$) polyhedra has been explored as a means to enhance birefringence in crystals, it is important to note that not all crystals containing such polyhedra exhibit exceptional birefringence. This is because the extent of birefringence enhancement is closely linked to the specific arrangement of these polyhedra. Therefore, a practical approach to exploring superior crystals is to use known structural templates as guidance. In the prototype structure templates, the contributions of birefringence-active polyhedra to the overall birefringence have been optimized. For instance, zero-dimensional (0D) $(\text{B}_3\text{O}_6)^{3-}$ units in $\alpha\text{-BBO}$ and two-dimensional (2D) $[\text{Be}_2\text{BO}_3\text{F}_2]_\infty$ layered frameworks in KBBF are known to represent the optimal templates for generating significant birefringence. Following this guidance, we proposed to use isolated $[\text{Sb-O-X}]$ polyhedra with larger anisotropy to resemble the $(\text{B}_3\text{O}_6)^{3-}$ units and to use $[\text{Sb-O-X}]$ layers to resemble the $[\text{Be}_2\text{BO}_3\text{F}_2]_\infty$ layers in quest of large birefringence. In these systems, the halogen ions and the cations with lone pairs usually act like “chemical scissors”, leading to the formation of low-dimensional frameworks that are favorable for generating large birefringence.⁶⁹

In light of the considerations mentioned above, the antimony(III) oxyhalide, Sb_2OCl_4 , was successfully designed and synthesized through the slow evaporation of a solution under open ambient conditions. Significantly, this compound features isolated $[\text{Sb}_4\text{O}_2\text{Cl}_8]$ units and demonstrates substantial birefringence. Notably, this finding underscores the significant contribution of arrangement with the isolated strong anisotropic $[\text{Sb-O-Cl}]$ polyhedron. Furthermore, in the context of the antimony-oxyhalide system, several compounds have been selected and investigated, including layered structures: $\text{Sb}_3\text{O}_4\text{F}$, $\text{Sb}_3\text{O}_4\text{I}$, $\text{Sb}_3\text{O}_4\text{Cl}$, $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and $\text{Sb}_8\text{O}_{11}\text{Br}_2$,^{70–77} and their birefringence all exhibit greater than 0.1.

In summary, this research encompassed the synthesis, UV-vis-NIR diffuse reflectance spectroscopy, IR spectroscopy, and thermal stability analysis of the compounds. First-principles calculations were also conducted for these compounds to elu-

cidate the relationship between their electronic structure and optical properties.

Results and discussion

Single-site substitution of the anions

Single-site substitution stands as an effective strategy in the synthesis of the crystals. In this approach, a single cation or anion within the crystal structure is replaced by another ion from the same group in the periodic table, typically when there is similarity in cation radii. This substitution can involve ions like NH_4^+ , K^+ , Rb^+ , and Cs^+ , leading to the creation of a wide range of exceptional materials. Anions are also amenable to single-site substitution, especially halogens. Nevertheless, the impact of such substitution on birefringence, such as when Cl atoms are replaced by the larger and more polarizable Br atoms, remains a subject of ongoing investigation. Inspired by the successful replacement of oxygen atoms in borate compounds with fluorine atoms to produce superior fluorooxoborates, we considered the replacement of the $\text{O}(\text{Cl})$ atom in SbOCl with the $\text{Cl}(\text{O})$ atom. This simple crystal served as a foundational template to explore the development of materials. As Fig. 1 shows, based on the foundational compound SbOCl , we designed and researched these antimony(III) oxyhalides: Sb_2OCl_4 , $\text{Sb}_3\text{O}_4\text{F}$, $\text{Sb}_3\text{O}_4\text{I}$, $\text{Sb}_3\text{O}_4\text{Cl}$, $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and $\text{Sb}_8\text{O}_{11}\text{Br}_2$, primarily through the substitution of single-site anions.

Structure description

All of the SbOCl , Sb_2OCl_4 , and $\text{Sb}_3\text{O}_4\text{F}$ compounds crystallize in space group $P2_1/c$. On the other hand, both $\text{Sb}_8\text{O}_{11}\text{Cl}_2$ and

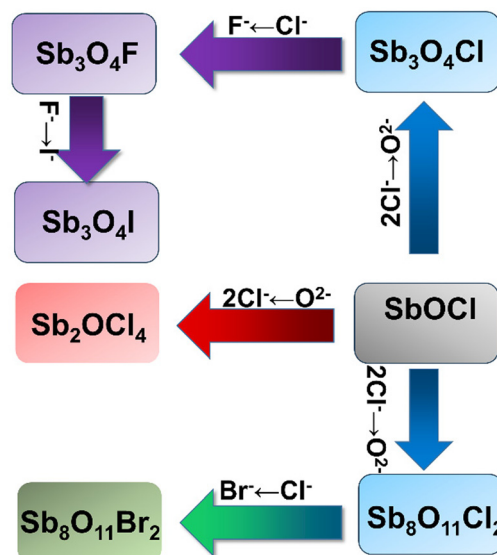


Fig. 1 Schematic illustration of the relationship between the template compound SbOCl and six antimony(III) oxyhalides, Sb_2OCl_4 , $\text{Sb}_3\text{O}_4\text{F}$, $\text{Sb}_3\text{O}_4\text{I}$, $\text{Sb}_3\text{O}_4\text{Cl}$, $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and $\text{Sb}_8\text{O}_{11}\text{Br}_2$, explored via single-site substitution.

$\text{Sb}_8\text{O}_{11}\text{Br}_2$ crystallize in space group $P\bar{1}$. However, $\text{Sb}_3\text{O}_4\text{Cl}$ and $\text{Sb}_3\text{O}_4\text{I}$ exhibit different space groups, specifically $P2/c$ and $Pna2_1$, respectively. It is worth noting that the bond lengths for the covalent Sb–Cl bonds in SbCl_3 range from 2.340 to 2.368 Å, while the Sb–Br bond lengths in $\alpha\text{-SbBr}_3$ and $\beta\text{-SbBr}_3$ range from 2.459 to 2.542 Å. In the case of $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, the shortest Sb–Cl bond length is 2.95 Å, and in $\text{Sb}_8\text{O}_{11}\text{Br}_2$, it is 3.14 Å. These relatively longer bond lengths, along with the small bond valence sum for the halide ions, suggest an ionic character of the Sb–X bonds in $\text{Sb}_8\text{O}_{11}\text{Cl}_2$ and $\text{Sb}_8\text{O}_{11}\text{Br}_2$.

The structures of four typical antimony oxyhalide compounds are described in Fig. 2a–d. As shown in Fig. 2a, each asymmetric unit of SbOCl comprises three Sb atoms, three O atoms, and three Cl atoms. The structure is composed of $(\text{SbO}_4)^{5-}$ and $(\text{SbO}_2\text{Cl})^{2-}$ groups. $(\text{Sb}_6\text{O}_7\text{Cl}_3)^+$ units are interconnected *via* oxygen atoms to form a channel structure. Chlorine ions are inserted into the channels to balance the charges, ultimately creating a three-dimensional structure.

As shown in Fig. 2b, each asymmetric unit of Sb_2OCl_4 consists of two Sb atoms, one O atom, and three Cl atoms. The crystal structure is built from $(\text{SbOCl}_3)^{2-}$ and $(\text{SbO}_2\text{Cl}_2)^{3-}$ groups. Isolated $[\text{Sb}_4\text{O}_2\text{Cl}_8]$ units are interconnected to form a three-dimensional structure. As shown in Fig. 2c, each asymmetric unit of $\text{Sb}_3\text{O}_4\text{Cl}$ includes two Sb atoms, one O atom, and two Cl atoms. The crystal structure forms a two-dimensional (2D) $[\text{Sb}_3\text{O}_4]_\infty$ layered framework, composed of $(\text{SbO}_4)^{5-}$ and $(\text{SbO}_3)^{3-}$ groups. Chlorine anions occupy the spaces between the layers to maintain charge balance. As shown in Fig. 2d, each asymmetric unit of $\text{Sb}_8\text{O}_{11}\text{Cl}_2$ includes sixteen Sb atoms, twenty-two O atoms, and four Cl atoms. The crystal

structure forms a 2D $[\text{Sb}_8\text{O}_{11}]_\infty$ layered framework, composed of $(\text{SbO}_4)^{5-}$ and $(\text{SbO}_3)^{3-}$ groups. Similar to $\text{Sb}_3\text{O}_4\text{Cl}$, chlorine anions are positioned between the layers to maintain charge balance.

Powder X-ray diffraction

The experimental X-ray Diffraction (XRD) curves closely match the theoretical ones for five compounds, as depicted in Fig. S1.† These results not only affirm the phase purity of the materials but also confirm the presence of each constituent element within the crystals, thereby substantiating the structural validity.

Optical properties

In Fig. S2,† the IR absorption spectra of SbOCl , Sb_2OCl_4 , $\text{Sb}_3\text{O}_4\text{I}$, $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and $\text{Sb}_8\text{O}_{11}\text{Br}_2$ reveal that the regions without obvious absorption for these powders fall within the range of 4000–725, 703, 705, 783, and 691 cm^{-1} , respectively. The IR absorption is induced by the Sb–O stretching vibrations. Fig. S3† illustrates the UV-vis-NIR diffuse-reflectance spectra for these compounds, which provide insight into their bandgaps. The bandgaps obtained by conversion are 4.02 eV for SbOCl , 3.83 eV for Sb_2OCl_4 , 3.20 eV for $\text{Sb}_3\text{O}_4\text{I}$, 3.59 eV for $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and 3.47 eV for $\text{Sb}_8\text{O}_{11}\text{Br}_2$. The compounds with a relatively large bandgap, which may have high laser damage thresholds, are suitable for practical applications. The analysis of element type by energy dispersive spectroscopy (EDS) confirms the existence of the Sb, O, and Cl elements of Sb_2OCl_4 (Fig. S4 in the ESI†).

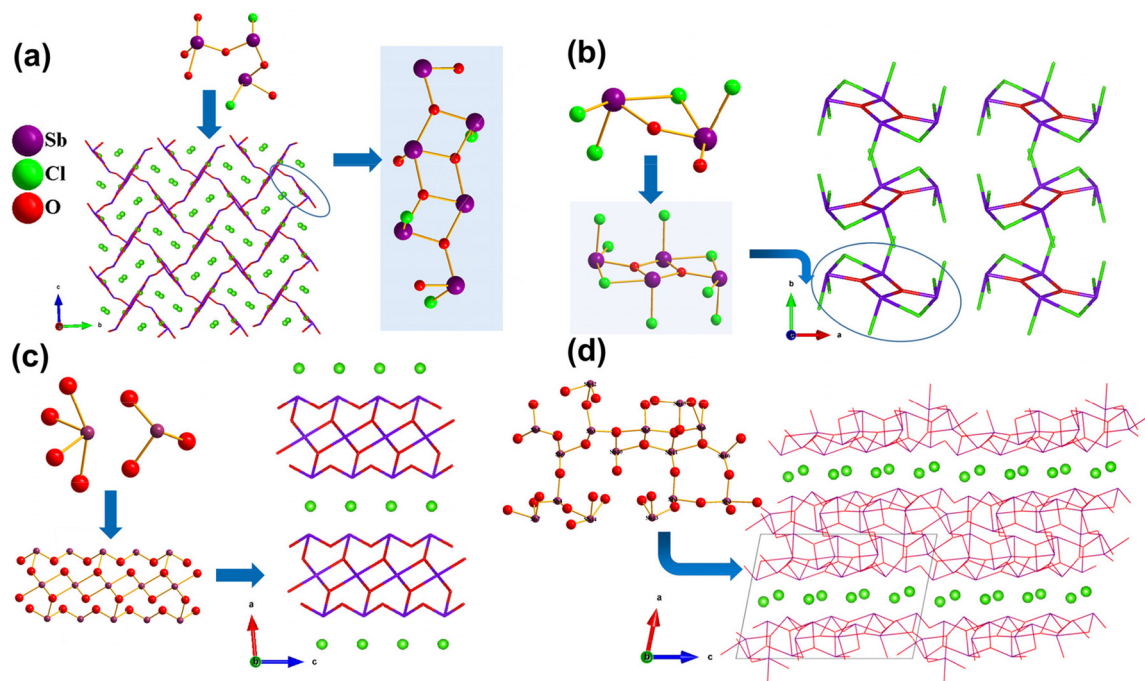


Fig. 2 Crystals structure information about compounds studied in this work. (a) The structure of SbOCl along the a axis; (b) the structure of Sb_2OCl_4 along the c axis; (c) the structure of $\text{Sb}_3\text{O}_4\text{Cl}$ along the b axis; (d) the structure of $\text{Sb}_8\text{O}_{11}\text{Cl}_2$ along the c axis.

Thermal stability

The thermal behaviors of SbOCl , Sb_2OCl_4 , $\text{Sb}_3\text{O}_4\text{I}$, $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and $\text{Sb}_8\text{O}_{11}\text{Br}_2$ are depicted in Fig. S5.† SbOCl , Sb_2OCl_4 , and $\text{Sb}_3\text{O}_4\text{I}$ exhibit stability up to approximately 200 °C, while $\text{Sb}_8\text{O}_{11}\text{Cl}_2$ and $\text{Sb}_8\text{O}_{11}\text{Br}_2$ remain stable up to approximately 500 °C. Differential scanning calorimetry (DSC) curves reveal several endothermic peaks for SbOCl , Sb_2OCl_4 , $\text{Sb}_3\text{O}_4\text{I}$, $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and $\text{Sb}_8\text{O}_{11}\text{Br}_2$, which implies that they break down. During the heating process, an evident weight loss is observed, and an endothermic peak emerges at around 220 °C, signifying a decomposition reaction in SbOCl . Upon further heating to 300 °C, the formation of liquid substances with a low melting point is evident in the product of the decomposition reaction. Simultaneously, an analysis of the melted samples indicates the presence of $\text{Sb}_4\text{O}_5\text{Cl}_2$ as a significant component, along with amorphous SbCl_3 . This suggests that the reaction can be described as follows:



Theoretical calculations

The electronic structure and optical properties of SbOCl , Sb_2OCl_4 , $\text{Sb}_3\text{O}_4\text{F}$, $\text{Sb}_3\text{O}_4\text{Cl}$, $\text{Sb}_3\text{O}_4\text{I}$, $\text{Sb}_8\text{O}_{11}\text{Cl}_2$, and $\text{Sb}_8\text{O}_{11}\text{Br}_2$ were investigated using first-principles calculations. The results along the high-symmetry points of the first Brillouin zone are depicted in Fig. S6.† It is worth noting that the calculated bandgaps for these compounds are smaller than the experimental values. This outcome aligns with expectations based on the density functional theory (DFT) method, which typically underestimates bandgap. This discrepancy is primarily attributed to inaccuracies in calculating the exchange–correlation energy. To address this issue, the Perdew–Burke–Ernzerhof (PBE0) hybrid functional was employed to obtain more accurate bandgap values. The difference between the bandgap value calculated by generalized gradient approximation (GGA) and the one by PBE0 (or the experimental value) was used as an adjustment factor to calculate the optical properties of these compounds. This approach helps refine the accuracy of the calculated optical properties, compensating for the initial underestimation of bandgaps by the DFT method.

Indeed, the behavior of valence electrons is widely recognized to have a crucial influence on the optical properties of a compound, particularly in the context of band structures near the Fermi level. Therefore, a more comprehensive examination of the electronic structure within this energy range was conducted to gain insights into its impact on bandgap and optical properties. As depicted in Fig. S6,† the electronic structures, density of states (DOS), and partial density of states (PDOS) are presented for the Sb, O, and X (X = halogens) orbitals in the seven compounds. This detailed analysis provides a deeper understanding of the distribution of electronic states and the contribution of specific atomic orbitals to the electronic structure, which, in turn, influences the optical properties and bandgaps of these materials.

Origin of the differences between the birefringence of compounds

The relationship between crystal structures and their corresponding microscopic groups is undeniably intertwined with the optical properties of materials. Factors such as the arrangement of anionic frameworks and the various combinations of metal cation polyhedra significantly influence properties like birefringence. To evaluate the birefringence, a polarizing microscope method was employed to test the difference in refractive index utilizing the natural growth plane. In Fig. 3a, the original interference color of the selected Sb_2OCl_4 crystal is shown under orthogonally polarized light. The crystal thickness is approximately 9.82 μm (Fig. 3b). As shown in Fig. 3c, the drum wheel of the Berek compensator was rotated to cause the crystal achieve extinction, and the optical path differences at a wavelength of 546 nm amount to 1.45 μm according to the tables of the Berek compensator specification. By applying the relevant formula, the refractive index differences were calculated to be 0.148. The experimental birefringences of Sb_2OCl_4 crystals were found to be greater than 0.148 at 546 nm, thus validating the calculated value.

Table 1 reveals that the Sb_2OCl_4 compound exhibits a birefringence of 0.175 at 546 nm, which is about 14.6 times that of MgF_2 . The improvement of the crystal structure of SbOCl through single-site substitution, along with well-considered halogen replacements, results in the creation of enhanced compounds, namely Sb_2OCl_4 , which exhibits an even more significant increase at 2.6 times the birefringence of the original compound. This birefringence value is on par with other recently reported materials, such as $\text{Sn}_2\text{B}_5\text{O}_9\text{Cl}$ (0.168 at 546 nm), $\text{Sn}_2\text{PO}_4\text{Cl}$ (0.181 at 546 nm), and SnF_2 (0.177 at 546 nm). These findings highlight the promising optical pro-

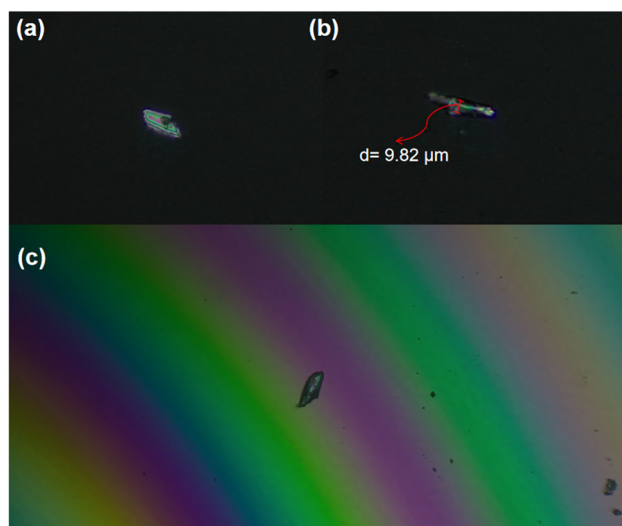


Fig. 3 Birefringence measurements for the Sb_2OCl_4 crystals. (a) Original interference color of the Sb_2OCl_4 crystal under orthogonally polarized light. (b) The thickness view of Sb_2OCl_4 crystal. (c) Extinction of the measurement of refractive index difference.

Table 1 The bandgap and birefringence of selected antimony halides, antimony oxyhalides, and commercial birefringent crystals²

Compound	Band gap (eV)	Birefringence
SbF ₃	4.30	0.104@546 nm ^a
Sb ₃ O ₄ F	3.82 ^a	0.192@546 nm ^a
SbOCl	4.02	0.067@546 nm ^a
Sb ₂ OCl ₄	3.83	0.175@546 nm ^a
Sb ₃ O ₄ Cl	3.69 ^a	0.107@546 nm ^a
SbCl ₃	4.51 ^a	0.173@546 nm ^a
Sb ₈ O ₁₁ Cl ₂	3.59	0.129@546 nm ^a
Sb ₈ O ₁₁ Br ₂	3.47	0.110@546 nm ^a
Sb ₃ O ₄ I	3.20	0.284@546 nm ^a
MgF ₂	11.27	0.014@193 nm
α-BBO	6.56	0.122@532 nm
CaCO ₃	3.54	0.171@633 nm
TiO ₂	3.10	0.256@1530 nm
YVO ₄	3.10	0.225@633 nm

^a Calculated value.

properties of Sb₂OCl₄ in comparison to other relevant compounds.

It is well-known that different structures and compositions often lead to distinct material performances. In explaining the differences in birefringence between SbF₃ and Sb₃O₄F, a comparative analysis of their structures and metal polyhedra is essential. To begin with, the Sb₃O₄F crystal possesses a monoclinic crystal system with lower symmetry than SbF₃. Additionally, the introduction of halogen atoms with stereochemically active lone pairs can enhance optical anisotropy. One key distinguishing factor is the coordination of metal polyhedra. The (SbO₃F)⁴⁻ polyhedra found in Sb₃O₄F have more oxygen ligands and exhibit a pronounced asymmetric coordination of metal polyhedra compared to the [SbF₃] polyhedra. The (SbO₃F)⁴⁻ polyhedron also demonstrates a more typical asymmetric structure of metal polyhedra due to the presence of lone pairs. These differences in structure and coordination contribute to the varying birefringence properties observed in Sb₃O₄F compared to SbF₃.

The arrangement of polyhedra with lone pairs profoundly impacts optical anisotropy. Using SbOCl as a template structure, structural modifications were made to create compounds with enhanced birefringence. In the case of Sb₂OCl₄, two chlorine (Cl) atoms were replaced with one oxygen (O) atom, leading to a significant gain in birefringence. Sb₂OCl₄ exhibits isolated [Sb₄O₂Cl₈] units, transforming the initially unfavorable pore structure into 0D birefringent active units reminiscent of α-BBO. On the other hand, by replacing one oxygen atom with two chlorine atoms, Sb₃O₄Cl was obtained. This compound features a distinct laminar structure, similar to KBBF, which contributes to the gain in birefringence. This structural modification, compared to SbOCl, leads to a significant increase in birefringence, emphasizing the role of structural transformations in enhancing optical properties.

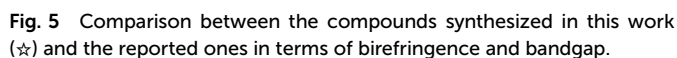
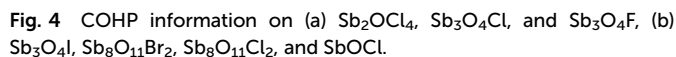
In alkali and alkaline-earth metal borates, it is noteworthy that many compounds with chlorine (Cl) are isostructural with

their bromine (Br) counterparts. However, the interesting observation is that there is often very little difference in birefringence between the chloride and bromide variants of these compounds. Examples of such compounds include K₃B₆O₁₀X (X = Cl, Br), Li₃B₈O₁₃X (X = Cl, Br), and Ba₃P₃O₁₀X (X = Cl, Br).^{78–84} This suggests that the alkali and alkaline-earth metal polyhedra have a relatively small influence on birefringence.⁹⁴ Thus, the change in halogen from Cl to Br does not significantly affect birefringence in these cases. However, the situation is different when the halogens are changed from chlorine to iodine. In these cases, there is a noticeable increase in birefringence. Examples of compounds exhibiting this trend include Pb₂BO₃X (X = Cl, Br, I), SnB₃O₇X (X = Cl, Br), SnB₅O₉X (X = Cl, Br), Sn₂PO₄X (X = F, Cl, Br, I), and PbSn(PO₄)X (X = Cl, Br, I).^{85–93} This indicates that the change in halogen from Cl to Br or I has a more pronounced effect on birefringence in these compounds, leading to increased optical anisotropy.

The introduction of halogens has a different effect on the birefringence of compounds containing cations with lone pairs, and this effect can be attributed to several factors. In the case of Sb₃O₄Cl and Sb₃O₄I, there is a gradual increase in birefringence from the chloride to the bromide compound. This trend is primarily related to the electronegativity of the halogen ions. This decrease in electronegativity strengthens the interaction between antimony and the halogen ions. As a result, the stereochemical activity of lone pairs in the [Sb–O–X] polyhedra increases gradually. This enhanced stereochemical activity of lone pairs is responsible for the observed increase in birefringence. Similar phenomena can be observed in other compounds, where changes in the halogen type (Cl⁻ to Br⁻) lead to an increase in birefringence. The combination of factors, including electronegativity and the arrangement of basic building units, contributes to the gradual increase in birefringence in compounds like Sb₃O₄X (X = Cl, I).

The absence of a noticeable increase in birefringence from Sb₈O₁₁Cl₂ to Sb₈O₁₁Br₂, despite the different polarization rates of chlorine and bromine, can be explained by examining the specific contributions of the [Sb–O–X] polyhedra and the Sb–Cl/Br bonds. In this case, the [Sb–O–X] polyhedra have comparatively weak Sb–Cl/Br bonds, and the differences in polarization rates between Cl⁻ and Br⁻ in these bonds have a limited impact on the overall birefringence. By contrast, the primary source of birefringence in these compounds comes from the [Sb–O] polyhedra. The larger atomic radius of bromine compared to chlorine does affect the density of the birefringent effective primitive, which can, in turn, lead to a smaller birefringence for Sb₈O₁₁Br₂ than for Sb₈O₁₁Cl₂. This effect is particularly relevant when considering the structural arrangement and contributions of the [Sb–O–X] polyhedra. This same phenomenon can be observed in Sb₃O₄X (X = F, Cl), where Sb₃O₄Cl has a smaller birefringence than Sb₃O₄F through Cl to F with a larger polarization rate. The primary contributing factor in these cases is the [Sb–O–X] polyhedra, while the Sb–Cl bond, due to its excessive length, does not significantly impact birefringence.

Our previous work sheds light on an interesting trend among oxyhalides based on the choice of cation.⁵¹ Oxyhalides with alkaline earth metals tend to exhibit large bandgaps but



All of the related experimental and computational data are provided in the ESI.†

Author contributions

C. H. H. and J. H. designed the research study; C. H. H. synthesized the compound; C. H. H. and X. L. H. performed the experiments. D. D. C. and F. Z. performed the optical theoretical calculations. All authors wrote and revised the manuscript. All the authors contributed to the final manuscript preparation.

Conflicts of interest

The authors declare no competing financial interests.

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