

REVIEW

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Emerging 15–16 group Xenes: structures, properties, preparation methods, and their catalytic applications

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Elemental two-dimensional (2D) materials, commonly referred to as Xenes, have attracted recent attention due to their many unique/remarkable chemical and physical properties. Xenes hold immense promise for multifarious applications across diverse domains, including optoelectronics, energy storage, energy conversion and biomedicine. Beyond graphene and phosphorene, a new cadre of Xenes has emerged, with particular attention directed toward antimonene, arsenene, tellurene and selenene. These nascent Xenes have garnered substantial interest due to their diverse allotropes, as well as their distinctive layer-dependent and modifiable properties, rendering them highly adaptable for engineering and catalytic applications. Herein, an overview is provided for the recent advancements in the structures, inherent properties and degradation behavior of Xenes, drawing upon both theoretical and experimental research. The synthesis methods of Xenes are summarized and primarily classified as bottom-up and top-down approaches. Furthermore, the catalytic potential of Xenes is elaborated, emphasizing both engineering strategies and theoretical understanding toward enhanced performance across a spectrum of catalytic reactions. Conclusively, a summary and perspectives on the future development of Xenes are given to boost their development.

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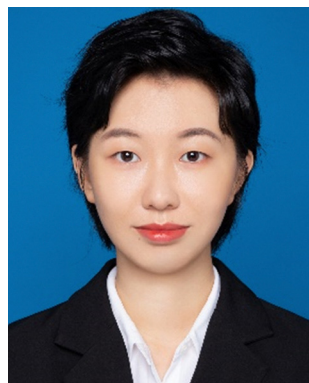
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1. Introduction

Two-dimensional (2D) materials are attracting extensive research interest across multiple disciplines, primarily due to their unique yet remarkable characteristics, including atomic thinness,^{1–3} large surface area,^{1,4,5} morphology tunability,^{4,6} structural anisotropy,^{7,8} excellent mobility^{4,9,10} and good mechanical flexibility,^{2,11,12} and the convenience to customize the surface characteristics for diverse applications.⁴ Among the vast 2D material family, elemental 2D materials belong to a developing branch of graphene-like 2D materials, commonly called “Xenes”.¹³ Xenes consist of one single element from the 13–17 groups in the periodic table, generally characterized by a layered architecture and large surface area similar to graphene,¹⁴ encompassing metals, semi-metals and non-metals.¹⁵ It is worth mentioning that α -phase and β -phase tellurene and selenene are intrinsically 1D materials in nature, composed of one-dimensional atomic chains. Therefore, they are classified as quasi-2D Xenes in this work,^{11,12,15} where a van der Waals interface between each layer can be observed and utilized in both chemical synthesis and exfoliation.

Xenes have many intriguing properties. One particularly notable and attractive aspect is their ability to modulate bandgaps through various surface engineering techniques, which contribute to multiple applications, including electronics, optoelectronics, sensors, energy storage, energy conversion, and biomedicine.¹⁴ Among Xenes, black phosphorene (BP) is particularly attractive and well-developed because of ultrahigh carrier mobility, tunable bandgap, and abundant active sites, which are favorable for catalysis. Nevertheless, the utility of BP is influenced by its susceptibility to degradation in water and air due to the vulnerability of its lone electron pair to ambient environments. Other elemental 2D materials within groups 15 and 16, such as arginine, antimonide, telluride, and selenite with BP-like marked puckered structures, have also attracted increasing attention in recent years due to their diverse crystal structures and properties and, more importantly, superior stability under ambient conditions.¹⁵

Several comprehensive reviews have been published on the Xenes family. Their myriad applications, including electronics,^{16–18} optoelectronics,^{13,16–18} energy storage,^{16–18} and conversion,^{16,17,19} as well as biological applications,^{16,20,21} picturing a vibrant research field. In addition, in-depth examinations are conducted for Xenes that attract extensive attention, such as phosphorene^{22,23} and borophene.^{24,25} Alternatively, this review focuses exclusively on the emerging research of Xenes from groups 15 and 16, including antimonene, arsenene, tellurene and selenene. These four emerging Xenes, located within groups 15 and 16, exhibit similar electron-structural properties owing to their adjacency within the periodic table, and exhibit tunable electronic band structures and high carrier mobilities^{16,18} and more importantly satisfactory stability compared to black phosphorous in the ambient environment.

Recently, there have been increasing studies on the roles of Xenes in clean energy applications, especially as electrocatalysts owing to the large specific surface area,^{4,5,17} high electron

transfer ability,^{26,27} tunable electron density^{28,29} and robust structural stability.^{16,30} However, as emerging 2D materials, the true potential of Xenes in the field of catalysis remains ambiguous. This review mainly discusses the emerging group 15 and 16 Xenes and their latest advancements in the past five years. Their commonalities and differences in structures, preparation methods, engineering approaches and catalytic applications are summarized and discussed as shown in Fig. 1. These insights offer a forward-looking perspective on the fundamental exploration and prospective application of Xenes' intriguing materials.

2. Layered structures and properties

Antimonene, arsenene, tellurene and selenene exhibit various phases, evident in both experimental synthesis and theoretical predictions. These Xenes can display semiconducting and metallic properties, as well as indirect–direct bandgap transitions, depending on their phase and layer number. Diverse atomic structures combined with the electronic characteristics of these emerging Xenes facilitate their extensive applications in energy conversion. In this review, we primarily focus on the phases that exhibit high chemical stability, and their structural properties are summarized in Table 1 and Fig. 2.

2.1 Group 15 Xenes

Group 15 elements, arsenic (As) and antimony (Sb), have electron configurations of $4s^24p^3$ and $5s^25p^3$, respectively, with five valence electrons. These elements require three additional electrons to achieve a stable octet configuration. A planar structure composed of sp^2 hybrids is unstable when it exists as a standalone layer. Hence, the introduction of sp^3 -like hybridization contributes to the stability of a buckled (β -phase) or puckered (α -phase) configuration (Fig. 3) for improved overlap of p_z - p_z orbitals.^{2,31,32} For the arsenene and antimonene monolayers, the sp^3 -like hybridization contributes to the formation of three σ covalent bonds with the adjacent atoms, while the lone pair of electrons in the p_z -like orbital interacts with atoms in neighboring layers, resulting in a pyramid of covalent bonds.³¹

2.1.1 Arsenene. Arsenene is notable for its allotropes, primarily gray arsenene (β -phase or buckled honeycomb arsenene) and black arsenene (α -phase or puckered honeycomb arsenene).^{21,31,38,55,56} Both can be obtained in nature, while other phases (γ , δ , ϵ , ξ , η , θ , ι) are revealed by limited theoretical works.⁵⁶ Compared to black phosphorus (BP), both gray and black arsenene exhibit better stability and possess cohesive energies similar to bulk arsenic.^{31,56} Notably, β -phase arsenene has a rhombohedral lattice structure with a space group of $R3m$, and its atoms follow an ABC stacking mode. It is recognized as the most stable configuration among all arsenene allotropes, which is critical for ensuring catalytic durability.^{57–60} The interlayer binding energy of gray arsenene is significantly stronger than that of graphene due to its cross-layer orbitals. This results in semimetallic properties and high



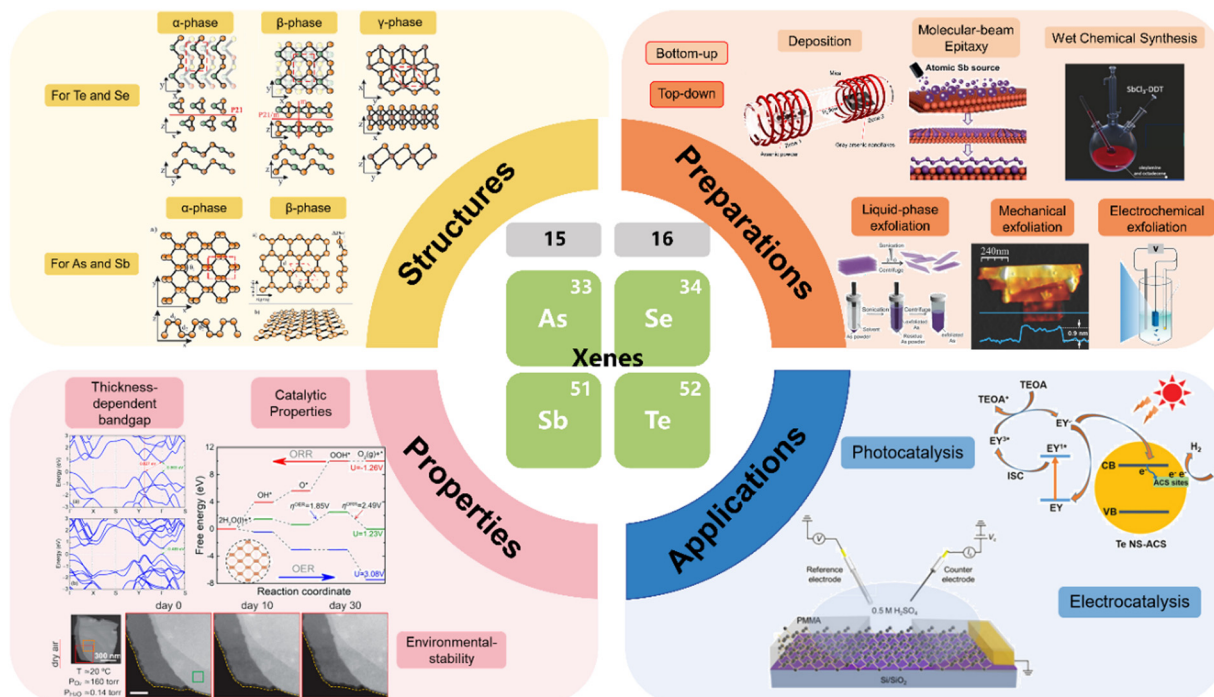


Fig. 1 Schematic of the review, illustrating the structures,³¹ preparation methods (deposition,³² molecular beam epitaxy,³³ wet-chemical synthesis,³⁴ liquid-phase exfoliation,² mechanical exfoliation³⁵ and electrochemical exfoliation²⁶), properties (thickness-dependent bandgap,³⁶ catalytic properties³⁷ and environmental-stability³⁸) and applications (photocatalysis³⁹ and electrocatalysis²⁷) of emerging group 15 and 16 Xenes. Reproduced from ref. 31 with permission from Royal Society of Chemistry, Copyright 2021; reproduced from ref. 32 with permission from American Chemical Society, Copyright 2019; reproduced from ref. 33 with permission from WILEY-VCH, Copyright 2019; reproduced from ref. 34 with permission from WILEY-VCH, Copyright 2019; reproduced from ref. 2 with permission from WILEY-VCH, Copyright 2019; reproduced from ref. 35 with permission from Wiley-VCH, Copyright 2016; reproduced from ref. 26 with permission from WILEY-VCH, Copyright 2021; reproduced from ref. 36 with permission from Springer Nature, Copyright 2015; reproduced from ref. 37 with permission from American Chemical Society, Copyright 2021; reproduced from ref. 38 with permission from American Chemical Society, Copyright 2020; reproduced from ref. 39 with permission from Wiley-VCH, Copyright 2023; reproduced from ref. 27 with permission from WILEY-VCH, Copyright 2020.

conductivity (10^2 – 10^4 $\text{cm}^2 \text{V}^{-1} \text{s}^{-1}$) for few-layer gray arsenene.^{2,40}

Black arsenene features an orthorhombic lattice similar to BP, characterized by a metastable phase with a space group of Bmab. Its synthesis typically necessitates precise kinetic control. Bulk black arsenic demonstrates semiconductive properties with a direct bandgap of 0.31 eV and exhibits a significant in-plane anisotropy along the armchair and zigzag directions.⁸ Notably, the properties of ultra-thin arsenene plates frequently diverge from those of bulk materials, highlighting the unique characteristics attainable at reduced dimensions.

Gray and black arsenene exhibit layer-dependent characteristics particularly in terms of their bandgap and dielectric response. For gray arsenene, the bandgap decreases upon increasing the layer number, varying from 0.37 eV (semimetal) of the bilayer to 2.49 eV (semiconductor) of the monolayer.^{2,55} Precious bandgap values exhibit slight variations depending on different predictive models. As the arsenene number increases from $n = 1$ to $n = 2$, black arsenene undergoes a transition from indirect-bandgap semiconductors (1–1.5 eV) to direct-bandgap metallic materials.^{36,37,61} The dielectric response of black arsenene, as measured by STEM-electron energy loss spectroscopy, shows a correlation with the number of layers.³⁸

Moreover, arsenic and arsenene can undergo a phase transition under specific conditions. For example, bulk arsenic undergoes a phase transition from black arsenic to gray arsenic at a critical pressure of 3.48 GPa, and transfers entirely to gray arsenic above 5.37 GPa. In the case of a few layers, semimetallic gray arsenene nanoflakes (4–48 nm) can be transformed into semiconductive vitreous arsenene nanoflakes through wet chemistry treatment *via* aqueous hydrofluoric acid (HF) solution.⁴⁰ Physical and chemical properties of antimonene, including mechanical strength, magnetic properties, electronic and optoelectronic properties, gas sensitivity, and catalytic activity, can be extensively tailored through various engineering methods, *i.e.*, doping,^{59,62} defect engineering,^{63–65} heterostructure formation,^{66,67} chemical absorption⁶⁸ and stain application.⁶⁹ For instance, theoretical investigations indicate that the introduction of dopants (Ti, V, Cr and Ni) effectively enhances the interaction between arsenene and hydrogen cyanide (HCN) gas, transforming the adsorption into chemisorption and thereby opening up new avenues for catalytic applications.⁶⁴

2.1.2 Antimonene. Compared to arsenene, antimonene is another pnictogen with higher thermodynamic stability. Although natural bulk antimonene is considered a half-metal,

Table 1 Comprehensive summary of antimonene, arsenene, selenene and tellurene, including structures, thickness, monolayer height, interlayer distance, lattice parameters, space group, bandgap, electronic properties and synthetic methods

Materials	Structures	Thickness (nm)	Monolayer height (Å)/interlayer distance (Å)	Lattice parameters (Å)	Electronic property/bandgap	Synthetic methods	Ref.
Vitreous arsenene	Amorphous (a regular hexagonal shape)	—	—	—	p-Type semiconducting (0.09–0.32 eV)	Wet chemistry vitrification	40
β-Arsenene	Rhombohedral (a regular hexagonal shape)	4–48	—	—	Semimetallic (pristine three-layer gray arsenene)	van der Waals epitaxy on mica	40
β-Arsenene	Rhombohedral	4	—	—	—	van der Waals epitaxial growth on mica	32
β-Arsenene monolayer	Buckled honeycomb	—	—	$a = b = 3.6$	Semiconducting	Grown on Ag (111)	41
β-Arsenene	Rhombohedral	5–12	1.35 (buckling height), 2.04 (inter-layer distance)	$a = b = 3.76$	—	Liquid-phase exfoliation (in organic and aqueous solvents using a tip sonication system)	2
α-Arsenene	Orthorhombic	≈0.6	—	—	—	Low-potential electrochemical exfoliation	42
β-Antimonene	Buckled honeycomb	—	1.78 (buckling height)	$a = 3.84$	0.11 eV (compressive-strained antimonene on Sb–Cu (110))	Grown on Cu (110) by MBE	33
β-Antimonene	Buckled honeycomb	—	1.42 (buckling height)	$a = 4.43$	0.21 eV (tensile-strained antimonene on Sb–Cu (111))	Grown on Cu (111) by MBE	33
β-Antimonene	—	≈4 nm in suspension, 2 μm thickness coating on TiO ₂ film	—	—	The hole acceptor layer in PEC	Liquid-phase exfoliation from bulk Sb crystal, and then using electrophoretic deposition on TiO ₂ films	43
β-Antimonene	—	1.71	—	—	1.57 eV (optical bandgap)	Li ⁺ intercalation and ultrasonication	44
β-Antimonene	—	4–17, average 10	9	—	—	A combination of ball milling and sonication-based liquid exfoliation	45
β-Antimonene	Rhombohedral	46	—	—	—	Soft-template assisted synthesis in colloidal solutions (SbCl ₃ dissolves in alkyl phosphonic acids)	46
α-Antimonene monolayer	Puckered structure	1–20 layers	≈0.78	$a = 4.4$, $b = 4.8$	Metallic	Grown on T _d -WTe ₂ by MBE	1
α-Antimonene monolayer	Puckered honeycomb	1–3 layers	—	$a = 4.46$, $b = 4.87$	Monolayer, bilayer and trilayer antimonene/MoTe ₂ are all metallic	Grown on T _d -MoTe ₂ by MBE	47
α-Antimonene monolayer	Puckered honeycomb	Stay α-phase to ≈15 layers	—	$a = 4.29$, $b = 4.76$	Semiconducting ≈0.17 eV (nearly free-standing monolayer), a semiconductor-semimetal crossover in bilayer α-Sb	Grown on SnSe by MBE	47
α-Selenene	Buckled non-planar	≈0.6–0.7	≈0.3	$a = b = 4.92$ and $c = 5.05$ (bulk)	Intrinsic p-type conductivity, ideal band-gap for visible light absorption (≈1.6 eV)	Grown on the silicon by CVT	48
α-Tellurene	Chiral-chain lattice stacked up by helical chains	10–50	—	—	—	Grown on exfoliated 1T'-MoTe ₂ by ultrasonication-assisted synthesis	9
α-Tellurene	Helical	Micrometer level	4.45 Å (distance between Te chains)	—	High hole mobility	Grown on 1T'-MoTe ₂ as the template by CVD	49
α-Tellurene	—	≈15	—	$a \approx b \approx 4.39$ Å	—	Grown on SrTiO ₃ (STO) by MBE	50
α-Tellurene	>100	—	—	—	Unipolar p-type transport	Grown on SiO ₂ /Si by CVD	39



Table 1 (continued)

Materials	Structures	Thickness (nm)	Monolayer height (Å)/interlayer distance (Å)	Lattice parameters (Å)	Electronic property/bandgap	Synthetic methods	Ref.
α -Tellurene	A helical chain	8–30	—	—	0.3 eV (bulk), 0.6 eV (4.5 nm-thick evaporated Te films) (optical bandgap), p-type mobility	Thermal evaporation at cryogenic temperatures	51
α -Tellurene	12–20	10–90	The interlayer spacing is ≈ 2.2 and 6, corresponding to the (1210) and (0001) planes.	—	Unipolar p-type transport	Hydrothermal	39
α -Tellurene	Bundles helical chains	≈ 27	The interplanar spacings 1.9 and 2.8, corresponding to Te (1210) and (0001) planes	—	Indirect bandgap (0.17 eV, bulk), indirect bandgap (0.27 eV, the synthesized Te), high hole mobilities	Hydrothermal	52
α -Tellurium	Three-fold screw-symmetric atomic chains	10–20	—	—	n-Type semiconducting	Hydrothermal	53
β -Tellurene	Helical chains with threefold screw symmetry	≈ 0.4 –0.8	Lattice spacings of ≈ 0.193 nm, and ≈ 0.225 nm, corresponding to the (200) and (110) planes	—	Indirect bandgap (bulk crystal), 0.09 eV (bulk) to 1.27 eV (monolayer), p-type semiconducting	Grown on WS ₂ by vdW epitaxy	54
γ -Tellurene	Hexagonal	Nanometer level	—	—	0.7 eV (monolayer)	PVD	3

β -antimonene and α -antimonene monolayers, two widely-reported allotropes with nonplanar structures, are theoretically predicted to be semiconductors with indirect bandgaps. Interestingly, metallic behavior is observed in epitaxial growth processes, which can be attributed to interactions with the substrate. Similar to arsenene, antimonene also exhibits various configurations, including three typical honeycomb (γ , δ , ϵ) and four non-honeycomb (ξ , η , θ , ι) structures.⁵⁶ β -phase is the most stable, featuring parallel buckled layers with the $R3m$ space group in a rhombohedral structure. In contrast, the α -phase consists of parallel puckered layers with the $Cmca$ space group. Despite its stability, antimonene possesses a tendency for oxidation,¹⁶ which is leveraged to optimize its electronic properties.⁷⁰

Recent reports reveal that various engineering approaches, such as interface engineering,^{33,47,71} defect engineering,^{72,73} doping,⁷⁴ strain application,^{75–77} and electron-beam irradiation,⁷⁸ play a vital role in tailoring the properties of antimonene. Interface engineering has proven effective in tuning the electronic structure of antimonene. For example, narrow-gap semiconducting α -antimonene monolayers are achieved through interface engineering. The specialized three-heights-buckling structural configuration of antimonene, controlled by interfacial interactions with substrate, significantly influences its electronic structure.⁷⁹ Interlayer strain-induced transitions between indirect and direct bandgaps, achievable through lattice mismatch induced by processes like epitaxial growth, enhance its potential for optoelectronic applications.¹ Under pressures exceeding 12%, bandgaps of antimonene decrease significantly, down to 0 eV.²⁹ Additionally, while introducing magnetism in antimonene, defects can impact geometrical symmetry and the band gap. External electric field and strain can easily modify antimonene electronic structure. For instance, Cr-doped antimonene can transform from a magnetic semiconductor to a magnetic semimetal under biaxial tensile strains of 6%.⁷⁵ Both Cr doping and biaxial compression strains contribute to the visible light absorption efficiency of antimonene.⁷⁵ Furthermore, electron-beam irradiation can alter the stacking structure of antimonene, with moiré patterns observed in twisted few-layered antimonene.⁷⁸

2.2 Group 16 Xenos

Chalcogenide atoms, such as selenene (Se) and tellurene (Te), each possess six valence electrons, with configurations of $4s^2 4p^4$ and $5s^2 5p^4$, respectively. To achieve a stable octet structure, Se and Te typically form two covalent bonds with their direct neighboring atoms, resulting in 1D-helical chains with a helical turn of 120° , which presents the most stable phase (α -phase). Another stable configuration is the tetragonal phase (β -phase), characterized by a mirror plane and rectangular unit cell, originating from a helix-to-layer transformation of the trigonal monolayer. Additionally, the trilayer rhombic structure with six bonds of the central atom and three bonds of the atoms in the outer layers is called γ -phase (1T-like phase).

Based on the transition from the nonmetal sulfur (S) to metal polonium (Po) within group VI, tellurene exhibits more



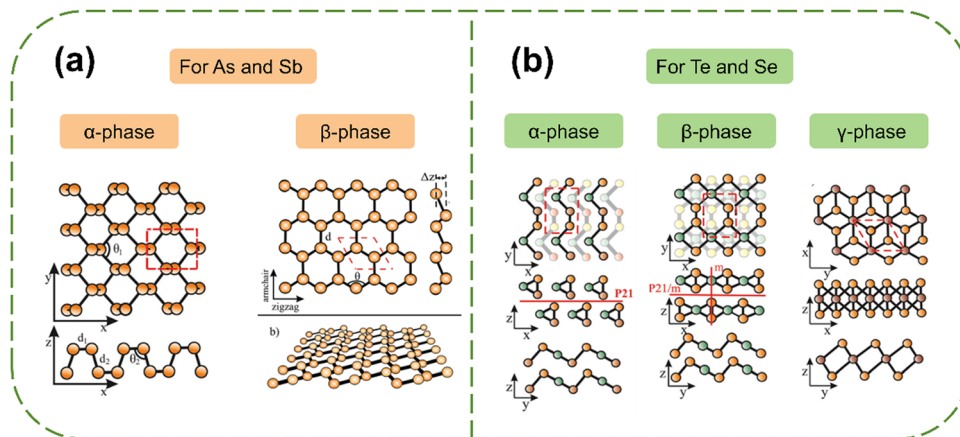


Fig. 2 Structures of antimonene, arsenene, tellurene and selenene. (a) α -phase and β -phase for arsenene or antimonene. (b) α -phase, β -phase and γ -phase for selenene and tellurene.³¹ Reproduced from ref. 31 with permission from Royal Society of Chemistry, Copyright 2021.

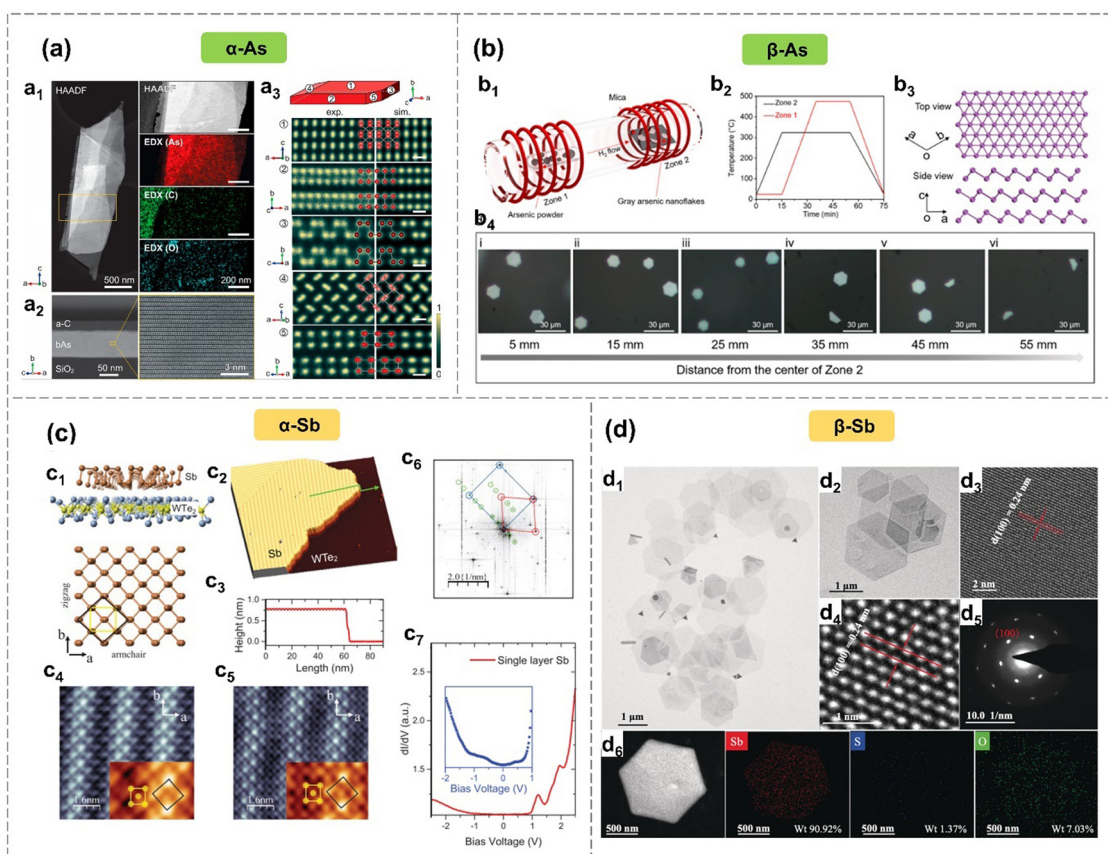
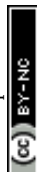


Fig. 3 Crystal structures of α -As,³² β -As³⁸ and α -Sb,¹ β -Sb.³⁴ (a) α -As prepared by mechanical exfoliation. HAADF-STEM and EDX elemental mapping images of α -As (a₁). Cross section of a β -As deposited on SiO₂/Si substrate (a₂). HAADF-STEM images of β -As from five orientations (a₃). (b) β -As synthesized by epitaxial growth on mica substrate. Schematic diagram of epitaxial growth of β -As (b₁). Temperature evolution curves of zone 1 and zone 2 during epitaxial growth (b₂). Buckled structure of β -As (b₃). Optical images of the synthesized β -As (b₄). (c) α -Sb epitaxial grown on T_d-WTe₂ substrate. The puckered honeycomb structure of α -Sb and WTe₂ (c₁). The STM topology image of monolayer α -Sb on WTe₂ (c₂) and the height variation along the green line (c₃). Atomically resolved STM images taken under $U = +300$ mV (c₄) and $U = -50$ mV (c₅), respectively. FFT image of the STM topology image (c₆). STS image of single-layer α -Sb (c₇). (d) β -Sb prepared via solution-phase synthesis. LRTEM images of β -Sb nanosheets (d₁, d₂). HRTEM images of β -Sb nanosheets corresponding to the (100) plane (d₃, d₄). SAED pattern along the [001] axis (d₅). HAADF-STEM and EDX elemental mapping images of β -Sb nanosheets (d₆). Reproduced with permission from: (a) ref. 32, with permission from American Chemical Society, Copyright 2019; (b) ref. 38, with permission from American Chemical Society, Copyright 2020; (c) ref. 1 with permission from Wiley-VCH, Copyright 2019; (d) ref. 34 with permission from Wiley-VCH, Copyright 2019.



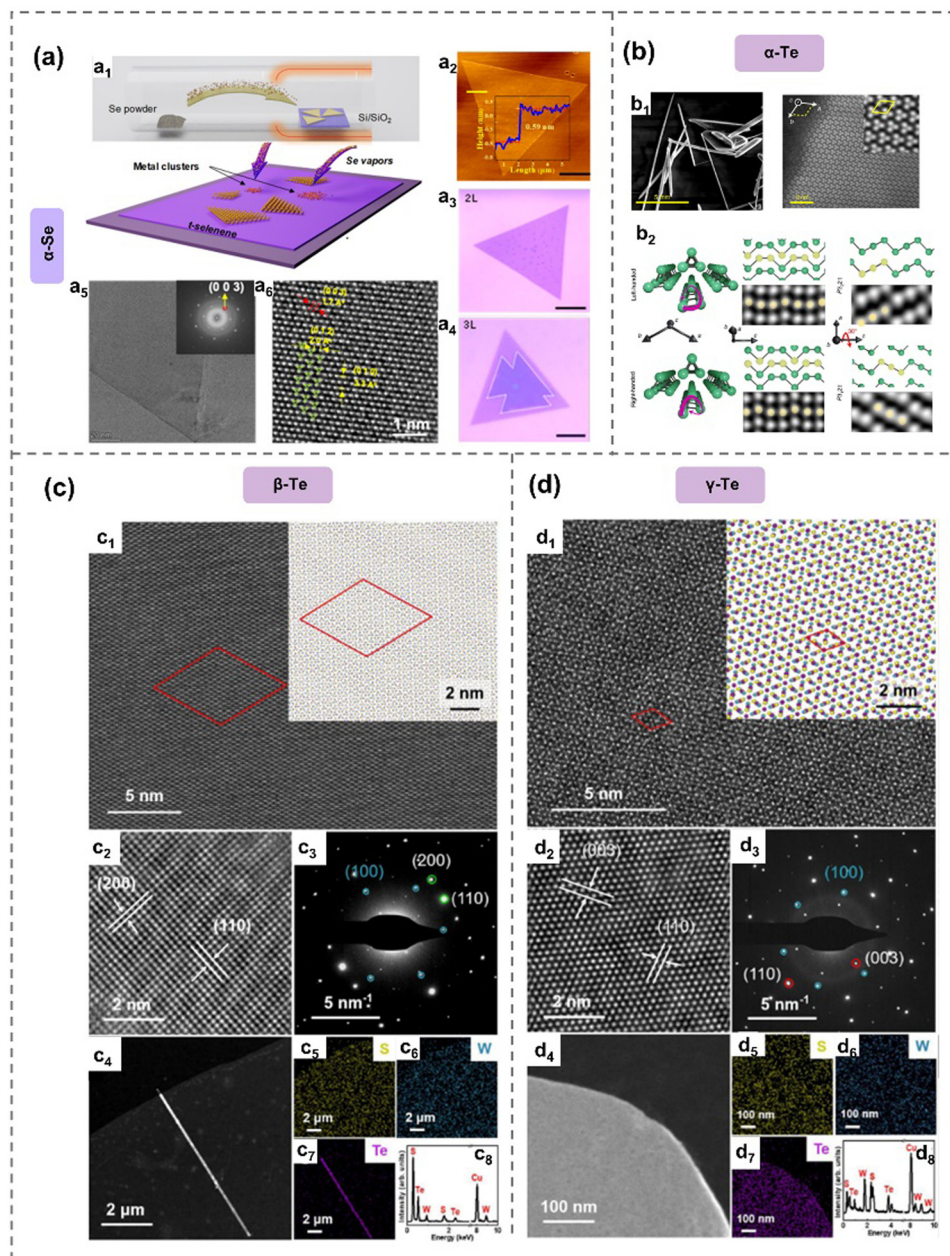


Fig. 4 Crystal structure of α -Se⁴⁸ and α -, β - and γ -Te.⁵⁴ (a) α -Se prepared by seed-assisted CVT. Schematic diagram of CVT process (a_1). AFM image of bilayer selenene (a_2). Optical images of bilayer selenene (a_3) and trilayer selenene (a_4). HRTEM image and SAED pattern (inset) of selenene domain (a_5). HAADF-STEM image of selenene. (b) α -Te nanowires prepared by hydrothermal method. SEM and STEM of α -Te (b_1). Crystal sketches and STEM images of α -Te (b_2). (c), (d) Crystal structure of β -Te and γ -Te prepared by vdW epitaxy method on WS_2 substrate. β -Te nanosheets (c_1) and γ -Te nanoribbons (d_1) on WS_2 . Insets display the β -Te/ WS_2 (c_1) and γ -Te/ WS_2 (d_1) heterostructures. HRTEM image of β -Te nanosheets (c_2) and γ -Te nanoribbons (d_2) on WS_2 . SAED patterns of β -Te/ WS_2 (c_3) and γ -Te/ WS_2 (d_3). TEM images of β -Te nanosheets (c_4) and γ -Te nanoribbons (d_4). EDS mapping and elemental analysis of α -Te nanosheets (c_5)–(c_8) and β -Te nanoribbons (d_5)–(d_8) on WS_2 . Reproduced from: (a) ref. 48 with permission from Royal Society of Chemistry, Copyright 2022; (b) ref. 87 with permission from Springer Nature, Copyright 2022; (c), (d), ref. 54 with permission from Springer Nature, Copyright 2024.

metallic properties. This is due to a stronger delocalization of electrons within the chains of tellurene, leading to covalent-like interactions that are stronger than the van der Waals interactions in selenene. In contrast, based on the calculated cohesive energy, α -Se (3.35 eV per atom) is found to be energetically more stable than α -Te (3.05 eV per atom).⁸⁰

2.2.1 Selenene. Theoretical predictions suggest three possible configurations of monolayer selenene. 1D helical-chain Se

(α -phase), tetragonal phase (β -phase) and 1T-MoS₂-like Se (γ -Se).⁸¹ α -, β - and γ -Se are identified as the stable phases of selenene, with α -phase (as shown in Fig. 4(a)) being successfully synthesized and extensively discussed.^{48,82} Tiled 1D helical-chain Se possesses a centrosymmetry-breaking structure, resulting in a spontaneous in-plane ferroelectric polarization of approximately $2.68 \times 10^{-10} \text{ C m}^{-1}$ per layer. α -Phase selenene showcases outstanding carrier mobilities on the order of



$10^3 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, while monolayer β -Se exhibits significant in-plane spontaneous polarizations.⁸¹ Additionally, electronic structure calculations reveal that both 1T-MoS₂-like Se and 1D helical chain Se exhibit indirect-band-gap semiconductor behavior, with a bandgap of 1.11 eV and 2.64 eV, respectively.⁸¹ Theoretical investigations indicate that selenene exhibits thickness-dependent optical properties.^{83,84} Furthermore, the electronic and thermoelectric properties of α -phase selenene can be effectively manipulated through impurity adsorption and substitutional atom introduction.⁸⁵ The introduction of defects significantly reduces the bandgap energy, leading to a semiconductor-to-metal transition.⁸⁵ Similar transition can also be achieved by thermal excitations, 2D compressive strain and the introduction of excess charge with positive and negative polarity.⁸⁶

2.2.2 Tellurene. Tellurene exhibits structural allotropes similar to selenene, with α -phase being considered the most stable phase, characterized by 1D parallel helical Te chains. α -Te (Fig. 4(b)) and γ -Te (Fig. 4(d)) phases exhibit trigonal crystal lattices with threefold and sixfold coordination structures, respectively, while the β -Te (Fig. 4(c)) phase displays a mixture of three- and fourfold coordination structures. Moreover, in the γ -phase, which has a hexagonal three-atomic-layer structure, the central Te layer exhibits more metallic behavior,⁸⁸ while the outer Te layers display more semiconducting-like characteristics. This results in the formation of metal-ligand type chemical bonds between Te layers.⁸⁸ 2D tellurene is reported to exhibit a trigonal crystal lattice, where Te atoms are arranged in a unique 1D chiral-chain lattice.^{9,89,90} These layers are stacked together through van der Waals-like interchain bonds and spiral around axes parallel to the [0001] direction, forming a hexagonal array with a space group of $P3_221$ or $P3_21$.⁹ Therefore, tellurene is recognized as a quasi-2D material whose structural arrangement impacts its inherent anisotropy and induces chirality-dependent properties.¹⁸ Tellurene band structure and carrier dynamics are probed by time-resolved polarized reflectance spectroscopy.⁹¹

2D tellurene demonstrates thickness-dependent bandgaps, varying from nearly 0.31 eV for bulk to 1.17 eV for bilayer.⁹² This material exhibits significantly enhanced optical absorption properties due to a nearly direct or direct band gap and high hole carrier mobility, making it a promising candidate for applications in (opto-) electronics especially the general scarcity of p-type 2D semiconductors.⁹³ Similar to selenene, the properties of tellurene, including the phase transition, bandgap and (opto-) electronic, magnetic, mechanic and catalytic properties, can be optimized through strain,^{94–99} surface adsorption,¹⁰⁰ vacancy,^{6,100} doping,⁹⁸ thermal excitations,¹⁰¹ electric fields manipulation¹⁰² and heterostructure formation.^{28,103}

Theoretical reports predict several interesting properties of tellurene waiting for experimental validations. First-principles calculations indicate that monolayer tellurene can undergo a β -to α -phase transition under biaxial strains to 0.5%, accompanied by a substantial increase in high piezoelectric coefficient.⁹⁶ And in-plane compression can facilitate the α -to β -phase transition. And the growth of the tensile strain along

the chain direction beyond a critical value, $\approx 12\%$, may lead to the α -to γ -phase transition.¹⁰⁴ The tellurene bandgap can transition from indirect to direct when the vacancy concentration reaches 5.3%.¹⁰⁰

Moreover, transition metal doping and strain engineering can introduce magnetism into α - and β -tellurene while modifying their magnetic moments and electronic properties. Hybrid structures, such as α -Te/h-BN, exhibit enhanced carrier mobility and tunable electronic properties, with boron (B) vacancy driving the transition of Te/BN bilayers from semiconductors to half-metals. Notably, trigonal crystals, including their 2D nanosheets, can be transformed into various metallic structures through reversible metal-semiconductor transitions induced by electrostatic charging.²⁸ Additionally, photoinduced-shear stain in Te films disrupts the screw symmetry of the crystal, contributing to the dynamic band structure.⁹¹ Furthermore, tellurium nanofilms exhibit excellent thermoelectric performance, with a room-temperature power factor of $31.7 \mu\text{W cm}^{-1} \text{ K}^{-2}$ and a ZT value of 0.63.¹⁰⁵

2.3 Chemical stability of 2D Xenes

Although arsenene, antimonene, selenene and tellurene are generally stable under ambient environments, it is crucial to consider the influence of oxygen and water during their preparation and storage. Herein, examples of their oxidation and degradation behavior, as well as their advantages and disadvantages in the latest research reports are summarized.

The sp^3 -like hybrid electron structures of As and Sb both have a lone pair of non-bonding electron pairs that can interact with electrons donated by oxygen or water molecules, forming a dative bond.^{33,38} Specifically, it is observed that the presence of dissolved oxygen in aqueous HF solution significantly accelerates the vitrification process of arsenene nanoflakes (Fig. 5(c) and (d)). This is evident from the significant disparity in the rate of the surface color change of As crystals when exposed to HF in ambient air compared to N₂.⁴⁰ The reaction between As₂O₃, induced by dissolved oxygen, and HF aqueous plays a crucial role in the formation of As defects/vacancies and unsaturated As atoms, facilitating the transition to the vitreous arsenene.⁴⁰

Additionally, the impact of the ambient environment on gray arsenene is directly documented through optical microscopy images (Fig. 5(a) and (b)). Gray arsenic nanoflakes exhibit a transformation from metallic to semiconducting behavior after prolonged exposure to ambient conditions, attributed to the oxidation of arsenic in the air. This also indicates that the control of oxidation degree may be a novel strategy for bandgap modification. The oxidation of arsenic would happen after extended ambient exposure, which can be effectively inhibited by protective coatings that include polyethylene glycol (PEG), polyacrylonitrile (PAN), poly (methyl methacrylate) (PMMA), and polyvinylpyrrolidone (PVP). These protective coatings are employed to safeguard the sample against chemical degradation and phase transition for over 50 days under an ambient environment.³²



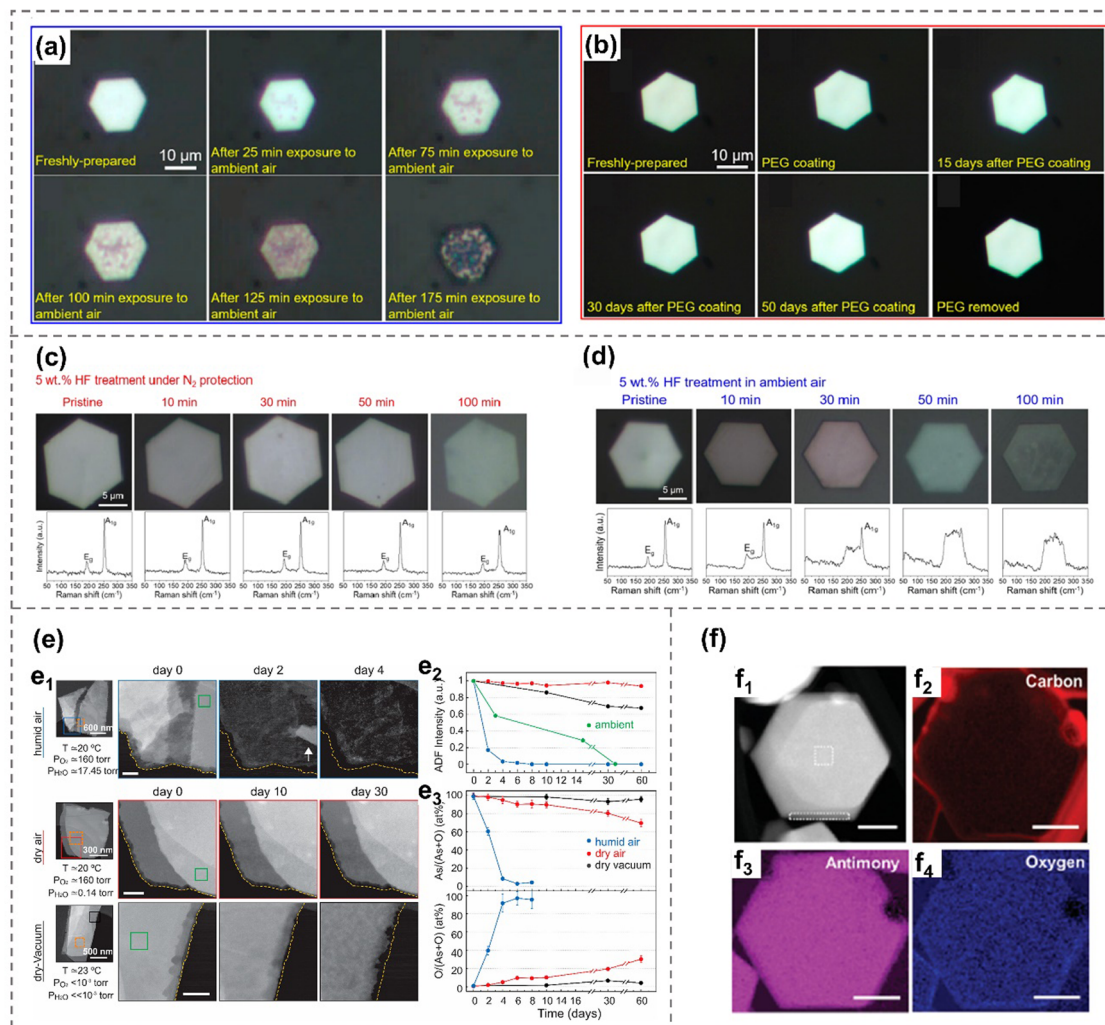


Fig. 5 Oxidation and degradation behavior of arsenene and antimonene. (a) Optical images of a β -arsenic nanoflake with varying durations of exposure to the ambient environment.³² (b) An as-prepared β -arsenene nanoflake, followed by its state after PEG coating and removal.³² (c) and (d) Optical images and corresponding Raman spectroscopy of β -arsenene nanoflake after different durations of HF treatment under N_2 protection (b) and in ambient air (c), respectively.⁴⁰ (e) Degradation of exfoliated α -arsenene flakes under humid air, dry air and dry-vacuum environment.³⁸ HADDF-STEM images showing degradation of α -arsenene over time (e_1). Changes in HAADF intensity of α -arsenene with time (e_2). Elemental composition of α -arsenene over time measured by EDX (e_3). (f) HADDF-STEM image (f_1) and elemental compositional EELS maps (f_2)–(f_4) of few-layer antimonene hexagons.¹⁰⁶ Reproduced from: (a), (b) ref. 32 with permission from American Chemistry Society, Copyright 2019; (c), (d), ref. 40 with permission from Wiley-VCH, Copyright 2021; (e) ref. 38 with permission from American Chemistry Society, Copyright 2020; (f) ref. 106 with permission from Wiley-VCH, Copyright 2021.

Besides, the oxygen and water degradation effect of black arsenene is examined *via* characterization techniques including optical microscopy, Raman, XRD, STEM and EDX spectroscopy.^{2,38} It shows that black arsenene exhibits a higher resistance to oxidation and vapor in wet air than gray arsenene,²¹ and the degradation rate of black As is more significant than that in dry air, and the oxygen content of the As flakes increases due to the vanishing of As in wet air. Yun *et al.* also reported that preventing exposure to water vapor is crucial for protecting black arsenic nanoflakes due to their heightened sensitivity to atmospheric humidity. Consequently, both dry air and a dry vacuum have been shown to be effective methods for safeguarding these nanoflakes.³⁸ In the case of a dry air condition, black As also undergoes an oxygen-increasing process even though there is no reduction of thickness (or As

content) and visible degradation, which is due to the formation of an As_2O_3 protection layer on the surface of black arsenene.³⁸

The oxidation phenomenon of Sb can be detected *via* the Sb $3d_{3/2}$ peak in the unannealed sample, as observed in XPS.³³ Additionally, it can also be detected by Raman modes centered at 190 cm^{-1} and 256 cm^{-1} .¹⁰⁰ This phenomenon can be further confirmed by EDX mapping images (Fig. 5(f)).¹⁰⁶ However, these oxidation layers can be effectively removed by annealing at 680 K, leading to high-quality antimonate without morphology changes.^{33,107} In addition to direct contact with chemicals and the ambient environment, electrochemistry can also induce oxidative behavior. For instance, Te can be readily oxidized to TeO_2 at merely 0.52 V (*versus* standard hydrogen electrode, SHE), which hinders its anodic exfoliation and restricts its application to cathodic exfoliation.¹⁰⁷

3. Preparation methods

The preparation methods of arsenene, antimonene, selenene and tellurene can be broadly categorized into two primary groups: top-down and bottom-up methods (Fig. 6). Mechanical exfoliation, liquid-phase exfoliation and electrochemical exfoliation are discussed within the scope of top-down strategies, while the bottom-up methods include molecular beam epitaxy, gas-phase deposition methods (including chemical vapor deposition, physical vapor deposition and electrophoretic deposition¹⁰⁸) and wet-chemical synthesis (thermal-liquid phase synthesis and thermal assistant techniques).

3.1 Top-down methods

3.1.1 Mechanical exfoliation. Mechanical exfoliation, although capable of generating high-quality single crystals with limited defects, suffers from the disadvantage of low yield and time-consuming process.^{108,109} This method is employed for the exfoliation of arsenene,⁸ antimonene³⁵ and tellurene¹⁰⁸ due to their layered or quasi-layered structures. For instance, bulk black arsenene prepared by chemical vapor transport technique is mechanically exfoliated using tape and subsequently transferred onto SiO₂/Si substrates *via* a polydimethylsiloxane (PDMS) stamp for TEM characterization.^{38,110} Bulk gray arsenic in nature can be exfoliated to obtain few-layer β -As.⁵³ Antimonene can also be mechanically exfoliated into thin films.¹¹¹

Due to the quasi-2D structure of selenene and tellurene, obtaining single-crystal nanosheets through mechanical exfoliation is challenging. Churchill *et al.* reported the mechanical

exfoliation of trigonal Te single crystals to produce nanoscale Te flakes and wires.¹⁰⁸ The silicon substrates require pre-treatment using acetone, isopropanol and O₂-plasma to enhance the adhesion of tellurene. The effect of directly sliding a cleaved facet of Te on the substrate is better than the scotch tape method, getting thicknesses of 1–2 nm and widths below 100 nm nanosheets.¹⁰⁸ To date, there are no reports on mechanical exfoliation of selenene to our best knowledge, this could be attributed to the scarcity of large, high-quality Se single crystals.¹⁰⁸

3.1.2 Electrochemical exfoliation. Electrochemical exfoliation stands out for its high efficiency in transforming bulk crystals into 2D layered Xenes. It involves applying an electric current to a precursor material immersed in an electrolyte, contributing to the separation of layers. In 2020, the bulk black arsenic crystal was the first to be exfoliated by low-potential electrochemistry in a non-aqueous, surfactant-free and oxygen-free *N,N*-dimethylformamide (DMF) medium assisted with ammonium hexafluorophosphate (NH₄PF₆).⁴² The as-exfoliated layered arsenene tends to display self-aggregation behavior during post-processing, manifesting as porous net-like agglomerates composed of various 2D flakes, each with a thickness of ≈ 0.6 nm and lateral areas in the micrometer range. A slight presence of As_xO_y on the surface is detected through Raman, XPS and EDX spectroscopy⁴² (Fig. 7).

Zheng *et al.* employed a cathodic exfoliation method to exfoliate tellurium crystals into Te nanosheets, achieved at an overpotential of approximately -0.4 V *versus* the reversible hydrogen electrode (RHE) in an alkaline solution (1.0 M

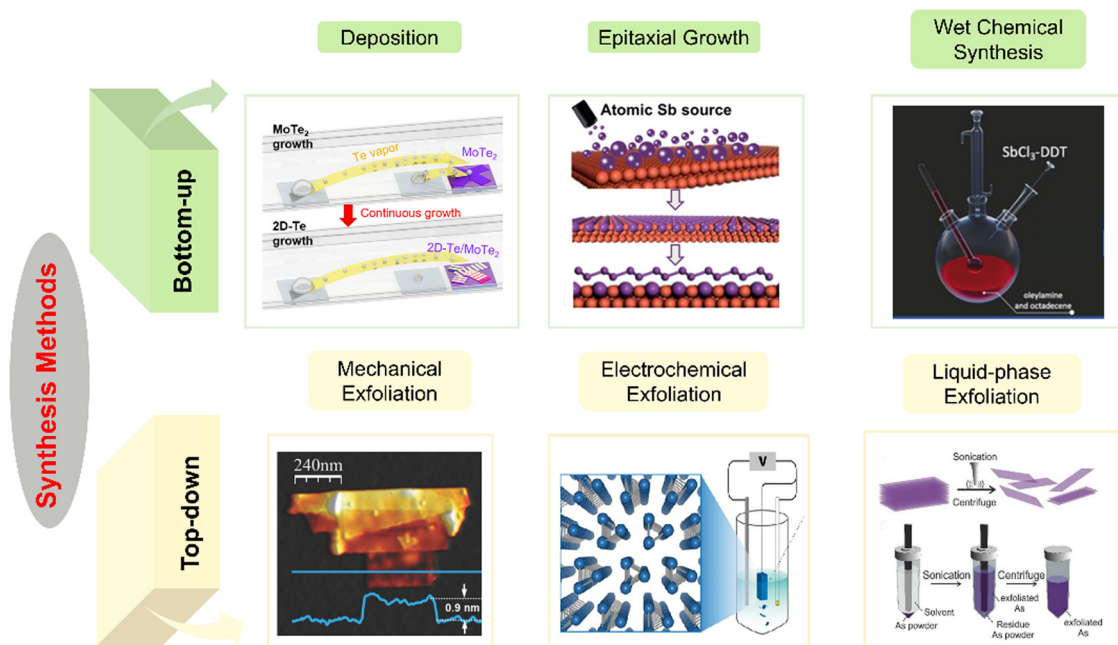


Fig. 6 Bottom-up and top-down synthesis methods of antimonene, arsenene, selenene and tellurene. Bottom-up methods include deposition,⁴⁹ molecular beam epitaxy³³ and wet-chemical synthesis.³⁴ Top-down methods include mechanical exfoliation,³⁵ electrochemical exfoliation²⁶ and liquid-phase exfoliation.² Reproduced with permission from: ref. 49 with permission from American Chemical Society, Copyright 2024; ref. 33 with permission from Wiley-VCH, Copyright 2019; ref. 34 with permission from Wiley-VCH, Copyright 2019; ref. 35 with permission from Wiley-VCH, Copyright 2016; ref. 26 with permission from Wiley-VCH, Copyright 2021; ref. 2 with permission from Wiley-VCH, Copyright 2019.

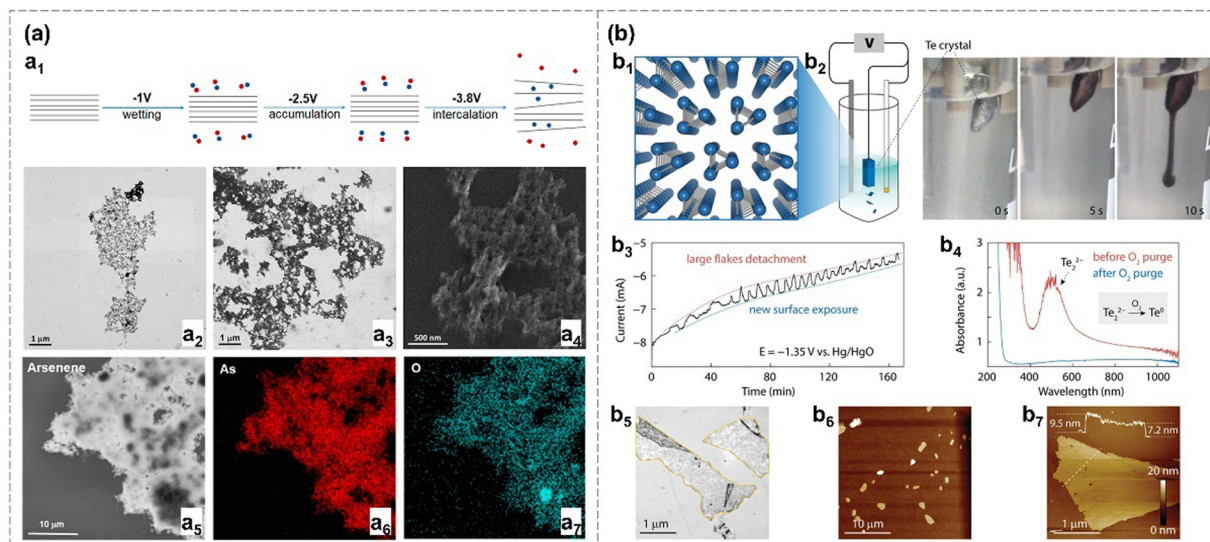


Fig. 7 (a) Low-potential electrochemical exfoliation.⁴² Schematic diagram of electrochemical exfoliation of arsenene. The blue ball represents cations (NH_4^+), while the red ball represents anions (PF_6^-) (a₁). STEM images of sonicated few-layer α -arsenene (a₂) and washed and sonicated α -arsenene (a₃), (a₄). EDX elemental mapping images of As and O element distribution (a₅)–(a₇). (b) Te crystal prepared by cathodic exfoliation.²⁶ Schematic diagram of Te crystal structure (b₁). Electrochemical exfoliation setup and digital photos of the exfoliation process over time (b₂). Amperometry curve of the exfoliation over time with a potential of -1.36 V vs. Hg/HgO (b₃). UV-vis spectroscopy of the exfoliated Te nanosheets before and after O₂ purge (b₄). Optical (b₅) and AFM images (b₆), (b₇) of Te on the Si/SiO₂ substrate. Reproduced from: (a) ref. 42 with permission from American Chemical Society, Copyright 2020; (b) ref. 26 with permission from Wiley-VCH, Copyright 2021.

KOH).²⁶ The reduction of Te^0 atoms at both the edge and plane regions to Te_2^{2-} facilitates the insertion of K^+ ions between the interlayers or interchains of Te, ultimately forming KTe_2^- . The KTe_2^- flakes with negative charge exhibit mutual repulsion, ensuring stability in solvent for 10 hours. These Te nanosheets can be collected with the assistance of dissolved oxygen, which prompts the conversion of Te_2^{2-} back to Te^0 . This electrochemical exfoliation method can contribute to the doping of metal in the active site, promoting the hydrogen evolution reaction (HER) activity of the materials.²⁶ Han *et al.* prepared high-quality antimonene nanosheets through an electrochemical lithiation intercalation process, followed by ultrasonication.⁴⁴ These Sb nanosheets serve as effective heterogeneous nucleation sites to promote the preferential growth of perovskite films.⁴⁴

3.1.3 Liquid-phase exfoliation. Liquid-phase exfoliation is a vital technique for isolating materials along van der Waals gaps within high-dimensional materials in a liquid environment to obtain 2D materials. This process involves dispersing bulk materials in a solvent, applying energy to separate the layers and using stabilizing solvents to prevent the reaggregation. This process enables the separation of samples with the desired thickness.¹³ Some common solvents that create stable dispersions are used including *N*-methyl-2-pyrrolidone (NMP), *N*-cyclohexyl-2-pyrrolidone (CHP), dimethylformamide (DMF), dimethyl sulfoxide (DMSO), and isopropyl alcohol (IPA). LPE is recognized as a potential method for the large-scale production of 2D materials,^{112,113} and it can also be used for the *in situ* functionalization of Xenex.¹¹⁴ However, this process may still induce defects to the materials.¹¹⁵ Due to the weak interlayer

bonding and brittle mechanical properties of arsenene and antimonene, shear exfoliation based on powder suspension stands out as an ideal mean of preparing 2D nanosheets.^{4,113} Ultrasonication-assisted liquid-phase exfoliation is effectively employed to produce high-quality, few-layer arsenene with precise control over sizes in solvents without any surfactants, water and oxygen involved.² And antimonene is reported to be stabilized when the solubility parameters of the solvent and solute are compatible.¹³

Additionally, *N*-methylpyrrolidone (NMP) has emerged as the preferred solvent, offering the highest concentration of stable arsenene sheets. Notably, the crystalline structure of arsenene, whether in zigzag or armchair directions, remains well-preserved during the LPE process (Fig. 8(a)). The exfoliated sample with a thickness of 5–12 nm is stable, although slight oxidation occurred, demonstrating the versatility of employing solvents with diverse boiling points and surface tensions for the exfoliation of arsenene. Exfoliated few-layer arsenene demonstrates promise in applications as gas sensors, which can selectively detect either methanol or acetone vapors depending on the chosen frequency.²

In the case of antimonene, the impact of initial crystal size, solvent selection and ultrasound parameters on liquid-phase exfoliation is explored. Wet-ball milling with 2-butanol, coupled with tip sonication yields superior results compared to bath sonication.⁷⁰ Additionally, a few-layer antimonene is prepared by a combination of ball milling and sonication-based liquid exfoliation in a 4 : 1 isopropanol/water mixture using Sb bulk crystals. The EDX spectroscopy shows the existence of oxygen on the prepared Sb nanosheets, which is consistent with

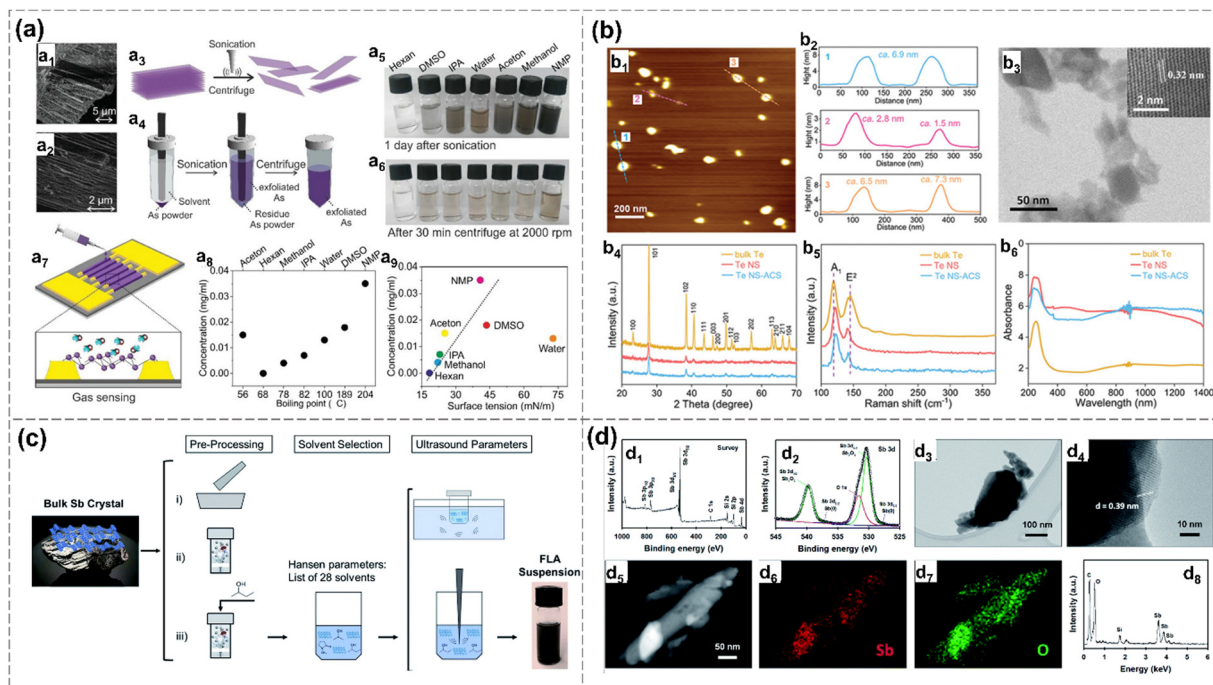


Fig. 8 Liquid-phase exfoliation methods. (a) Arsenene prepared by liquid-phase exfoliation and tip ultrasonication in different solvents.² SEM images of a layered bulk arsenic (a_1), (a_2). Schematic diagram of the tip ultrasonication and centrifugation process (a_3), (a_4). Digital photos of arsenene dispersed in different solvents after ultrasonication (a_5) and after centrifugation (a_6). Schematic diagram of vapor sensing setup based on prepared arsenene (a_7). Arsenene concentration in various solvents with different corresponding boiling point (a_8) and surface tension (a_9) after 2000 rpm centrifugation. (b) Atomic cobalt species functionalized Te nanosheets (Te NS-ACS) prepared by liquid-phase exfoliation chemical functionalization.²⁷ AFM image (b_1), thickness distribution (b_2), TEM and HRTEM images (b_3) of Te NS-ACS. XRD pattern, Raman spectroscopy and UV-vis absorption spectroscopy of bulk Te, Te NS and Te NS-ACS. (c) Optimization strategies of liquid phase exfoliation. Pre-processing is (i) grinding, (ii) dry ball-milling, (iii) wet ball-milling bulk Sb.¹¹² Solvent selection based on Hansen parameters using 28 different solvents (iii). Ultrasonication parameters comparison of bath and tip sonication. Method (iii) exhibits the optimum route to obtain FL Sb with the highest concentration. (d) FL-Sb nanosheets obtained by exfoliation using ball-milling and ultrasonication.⁴⁵ XPS (d_1), Sb 3d spectrum (d_2), TEM (d_3), HRTEM (d_4), HAADF-STEM (d_5), EDX elemental mapping (d_6), (d_7), EDX spectrum (d_8) of FL-Sb flakes. Reproduced from: (a) ref. 2 with permission from Wiley-VCH, Copyright XXXX; (b) ref. 27 with permission from Wiley-VCH, Copyright 2020. (c) ref. 112 with permission from Royal Chemistry of Society, Copyright 2019; (d) ref. 45 with permission from Royal Chemistry of Society, Copyright 2020.

XPS results (Fig. 8(d)).⁴⁵ Gibaja *et al.* applied standard LPE with three different preprocessing methods: grinding the bulk Sb crystal, dry ball-milling, and wet ball-milling (Fig. 8(c)). These methods do not have a significant impact on the in-plane structure, crystal phase, or composition of the exfoliated antimonene. The wet ball-milled samples exhibit the highest concentration.¹¹² Similarly, 2D tellurene is successfully exfoliated by LPE.^{27,116} Tellurene nanosheets (NS, 2–8.2 thickness) are exfoliated using the bulk tellurene in NMP solution *via* an ultrasonication device. The as-exfoliated samples react with $\text{CoCl}_2 \cdot 6\text{H}_2\text{O}$ to get the Te NS-ACS (atomic cobalt species), which is indicated by XRD, Raman and UV-vis absorption spectra (Fig. 8(b)).²⁷

3.2 Bottom-up methods

3.2.1 Gas-phase deposition. Abundant research is conducted on the fabrication of tellurene and selenene based on chemical vapor deposition (CVD), chemical vapor transport (CVT) and physical vapor deposition (PVD). Recent advancements demonstrate that low-temperature ($-80\text{ }^{\circ}\text{C}$) thermal evaporation is employed to obtain 2D trigonal tellurene films on various substrates, including silicon, glass and plastic for

the fabrication of wafer-scale field-effect transistors.⁵¹ The evaporated Te films exhibit distinct domains corresponding to arrays of aligned Te chains (Fig. 9(a)). The size of the domain can be effectively controlled by adjusting the temperature of the substrate. Furthermore, the optical properties of domains vary with thickness. SAED patterns of two regions with different contrasts reveal diffraction spots of single-crystalline structures with different orientations (Fig. 9(a)), indicating that both areas are single-crystal-like domains. HRTEM image clearly depicts the grain boundary of the two domains.⁵¹

Zhao *et al.* revealed that the grain size and orientation of the evaporated films can be finely controlled through crystallization processes on cold SiO_2/Si substrate ($-80\text{ }^{\circ}\text{C}$).⁸⁹ As-deposited Te film is amorphous and the transformation from amorphous to crystalline is shown in Fig. 9(b). New crystalline nuclei appear within the non-transformed portions of the sample throughout the transformation process. Eventually, the entire film undergoes a transformation into a polycrystalline structure, which is confirmed by the occurrence of the (100) peak from the XRD measurement (Fig. 9(b)).⁸⁹ The crystal growth kinetics related to nucleation rate, nucleus free energy, interface free energy and temperature is explained



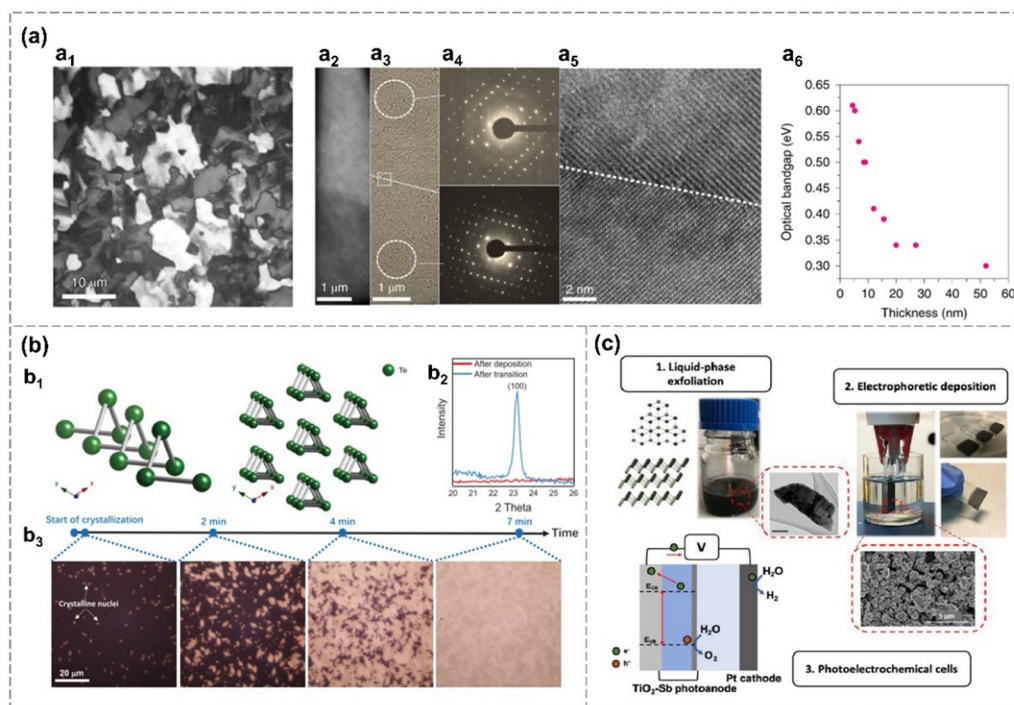


Fig. 9 Deposition methods. (a) Evaporated Te thin films on a $-80\text{ }^{\circ}\text{C}$ substrate.⁵¹ Polarized light microscopy of a Te film (a_1). Optical image of a Te film on SiO_2 TEM grid, showing a grain boundary (a_2). TEM images of a Te film (a_3) and corresponding SAED pattern (a_4). HRTEM image of the same film and the grain boundary (a_5). Optical bandgap changes over the thickness of a Te film. (b) Phase transition of evaporated Te films.⁸⁹ Te crystal structure (b_1). XRD pattern of a Te film (50 nm) evaporated on the $\text{SiO}_2/\text{p}^+\text{-Si}$ substrate ($-80\text{ }^{\circ}\text{C}$) after deposition and crystallization (b_2). Optical microscopy images of the crystallization process for an as-deposited 10 nm-thickness Te film (b_3) at room temperature. (c) Electrophoretic deposition (EPD) of liquid-phase exfoliated antimonene on conductive substrates and its activity as a light absorber and hole acceptor layer in photoelectrochemical cells (PECs).⁴³ Reproduced from: (a) ref. 51 with permission from Springer Nature, Copyright 2020; (b) ref. 89 with permission from Royal Chemistry of Society, Copyright 20201; (c) ref. 43 with permission from Elsevier, Copyright 2020.

theoretically.⁸⁹ Highly crystalline tellurium (Te) nanosheets, featuring well-defined shapes and sharp edges, are successfully synthesized through CVD.¹¹⁷ The growth mechanism is governed by surface energy considerations, with [001] and [100] orientations experiencing faster growth rates compared to other directions. These Te nanosheets demonstrate impressive capabilities in blackbody-sensitive infrared detection, exhibiting exceptional blackbody and polarization responses.¹¹⁷ Hydrogen is applied to reduce TeO_2 to synthesized Te nanosheets, with H_2O formation promoting the formation of Te flakes.⁹⁰ Te nanosheets are also successfully prepared *via* continuous template growth on $1\text{T}'\text{-MoTe}_2$ *via* CVD.⁴⁹

A seed-assisted CVT growth is employed to produce ultrathin, highly-crystalline α -Se nanosheets with a triangle-shape.⁴⁸ These nanosheets are oriented in (0001) direction and exhibit lateral dimensions exceeding $30\text{ }\mu\text{m}$ with an ultralow thickness of 0.59 nm (Fig. 4(a)). Sb, Bi and Sn are employed as seeding metals for the growth of α -Se. α -Se growth mechanism involves two steps: the initial formation of 3D selenium islands followed by lateral growth into a triangular shape. The nano-islands formed at both the edges and interiors during the initial growth phase are subsequently driven toward the edges of the domains with the progression of domain growth. The introduction of hydrogen atoms, serving as passivation agents for terminal Se

atoms, proves instrumental in stabilizing the lattice of α -Se nanoflakes.⁴⁸ α -Se flakes are reported to grow directly on the c -axis of the sapphire substrate *via* physical vapor deposition (PVD), contributing to the fabrication of α -Se based Schottky barrier diodes (SBDs).⁸² And hydrogen-assistant PVD is also utilized in the synthesis of tellurene. The introduction of hydrogen enhances the vapor pressure within the system and induces the reduction of TeO_2 , ultimately facilitating the formation of tellurene.⁴⁸ Additionally, as shown in Fig. 9(c), the LPE-prepared β -antimonene film can also be deposited on various substrates, such as a TiO_2 coated fluorine-doped tin oxide (FTO) electrode, through electrophoretic deposition, exhibiting good optical absorption and charge separation ability in the photoelectrochemical cell (PEC).

The synthesis of arsenene using the CVD technique faces certain limitations due to the vitality and thermal stability of the arsenic precursors. However, in industrial processes, arsenic precursors such as hydrides and alkyl/aryl hydrides are broadly employed due to their low thermal decomposition properties. This allows for the design of a suitable substrate for the future deposition of arsenene in large areas.¹¹⁸

3.2.2 Molecular beam epitaxy. Beyond the deposition method, molecular-beam epitaxy (MBE) plays a central role in the fabrication of Xenes. MBE synthesis of antimonene is



achieved on metallic, semi-metallic and dielectric substrates. The growth intermediates, phase transitions during the MBE process, as well as the effect of interlayer strain on the electronic properties of antimonene are discussed in detail. Scanning tunneling microscopy (STM)/scanning tunneling spectroscopy (STS) and computational methods are commonly used in the investigation of the growth process.

For instance, β -antimonene monolayers are successfully prepared on PdTe_2 substrates, which have a small lattice mismatch with that of free-standing Sb and high chemical stability.¹⁰⁷ Multilayer antimonene nanoribbon-like flakes, crystallizing in the β -phase and extending over several microns in length, are realized through molecular beam epitaxy (MBE) and subsequent *in situ* thermal annealing on sapphire substrates.¹¹⁹ The formation mechanism may be attributed to the transition of thermodynamic shape, enabled by the enhanced adatom mobility at elevated annealing temperatures.¹¹⁹ Shi *et al.* reported the successful large-scale synthesis of monolayer α -antimonene *via* MBE, employing Td-WTe_2 as the substrate because of its surface chemical inertness

and lattice matching with α -Sb.¹ A monolayer growth mode is initially favored, characterized by the formation of defect-free Sb islands with no detectable defects, and no phase transition during growth is detected.¹ Besides, single-crystalline β -antimonene is prepared by MBE on Cu (111) and Cu (110), respectively (Fig. 10(a) and (b)).³³ The antimony is evaporated under 640 K, and deposited on the copper surface at room temperature, forming twofold symmetry Cu-Sb surface alloy and threefold symmetric antimonene. The lattice mismatch between the epitaxial substrates and 2D antimony coverage can introduce compressive or tensile strain, contributing to various bandgaps of antimonene.³³ Additionally, an antimonene monolayer with a three-height-buckling configuration overlaid on a SbAg_2 surface alloy-covered Ag (111) substrate is synthesized by molecular beam epitaxy.⁷⁹

Besides metallic substrates, antimonene can also be epitaxially grown on dielectric substrates such as copper oxide through the MBE segregation process, which ranges from Sb dimer chains and hexagonally packed dot arrays. Eventually, Sb spreads over the surface and forms monolayer antimonene.

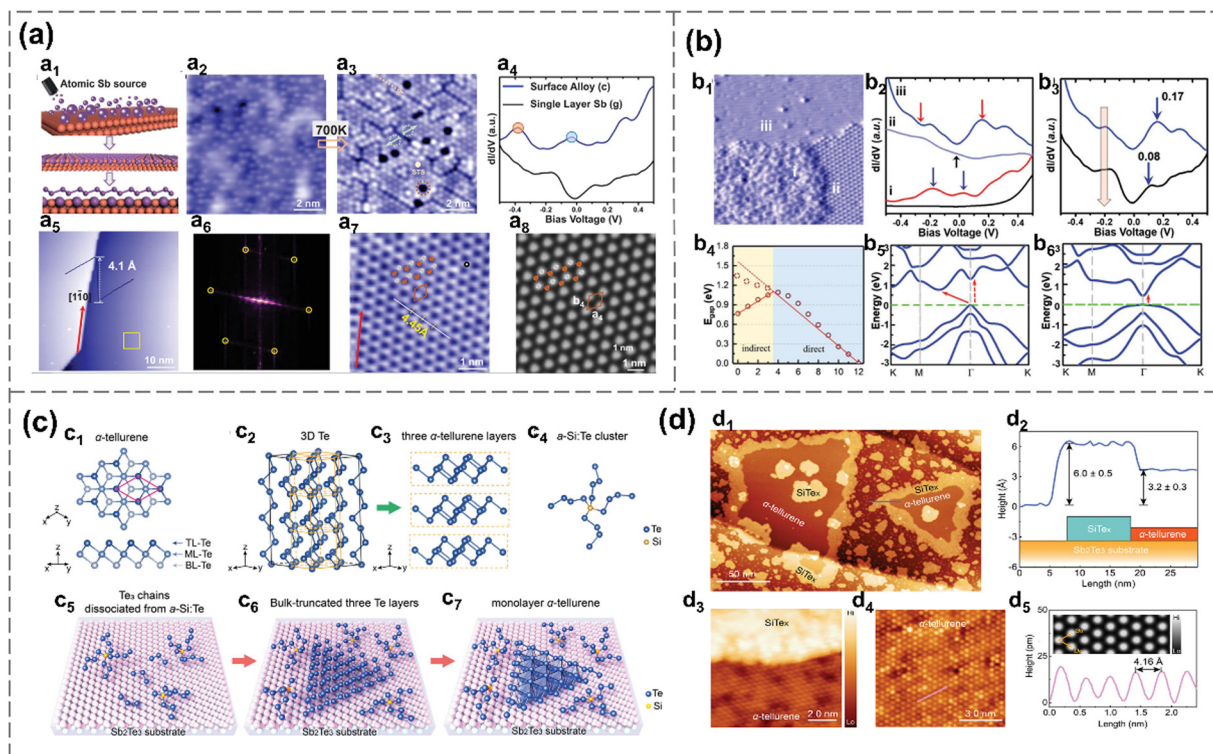


Fig. 10 Molecular beam epitaxial. (a) Crystal structure and electronic property of antimonene on Cu (111).³³ Schematic diagram of MBE process: dispersed Sb atoms, alloy with substrate and as-prepared antimonene (a₁). STM topographic image of 0.5 monolayer Sb on Cu (111) before annealing (a₂) and after annealing (a₃). STS of surface alloy and single layer Sb (a₄). STM image of well-ordered monolayer antimonene on Cu (111) after annealing at 700 K (a₅) and corresponding FFT (a₆). High-resolution STM image of antimonene taken from the yellow zone in a₅ (a₇). Simulated STM image of antimonene (a₈). (b) Electronic structure of antimonene grown on Cu₃O₂/Cu(111).²⁹ Grain boundary of antimonene.⁸⁸ (i) clusters, (ii) Sb pattern and (iii) antimonene (b₁). STS spectra of bare Cu₃O₂/Cu(111) and three phases as illustrated in b₁ (b₂). STS spectra of antimonene on Sb–Cu (111) alloy surface (b₃). Bandgaps of free-standing Sb under tensile strain calculated by DFT (b₄). Band structures of free-standing Sb (lattice constant is 4.01 Å) (b₅) and Sb on Cu₃O₂/Cu(111) (where lattice constant of Sb is 4.35 Å) (b₆). (c) Synthesis of γ -Te.⁸⁸ Schematic diagram of γ -Te monolayer (c₁), 3D Te chains (c₂), tri-layer Te (c₃), Te-rich a-Si:Te cluster (c₄). Epitaxial growth steps of Te on Sb₂Te₃ substrate (c₅–c₇). (d) Morphology of Te grown on a Sb₂Te₃ substrate.⁸⁸ STM topographic image of γ -Te grown on Sb₂Te₃ (d₁). Edge-step between γ -Te and SiTe_x (d₂). Atomic-resolution STM image of γ -Te (d₃). Height along the pink line in d₃ (d₄). Height along the blue line in d₁ (d₅). Reproduced from: (a) ref. 33 with permission from Wiley-VCH, Copyright 2019; (b) ref. 29 with permission from Wiley-VCH, Copyright 2020; (c), (d), ref. 88 with permission from Wiley-VCH, Copyright 2024.

There is a weak van der Waals force formation (as indicated by electronic localization function) between $\text{Cu}_3\text{O}_2/\text{Cu}(111)$ and antimonene, preserving its semiconductor property (direct bandgap, 0.37 eV).²⁹

Additionally, MBE is employed to cultivate high-performance optoelectronic nanodevices. Kang *et al.* synthesized 2D single-crystalline Te nanobelts on SrTiO_3 (001) substrates. This synthesis involves a preliminary substrate treatment consisting of combined chemical etching and thermal annealing. The process culminates in nanobelts that exhibit impressive photoconductive performance, including a broad spectral response range from the near-ultraviolet to the near-infrared regions, along with rapid response and decay times of 510 μs and 300 μs , respectively.⁵⁰

Moreover, a self-assembly approach, using $\alpha\text{-Si:Te}$ alloy as building blocks, is employed to grow γ -phase monolayer tellurene on a Sb_2Te_3 substrate epitaxially. First, twofold coordinated Te chains are linked together by fourfold coordinated Si atoms to form an alloy (Fig. 10(c) and (d)). Then, under high temperatures, the structure of the alloy softens, and the twofold coordinated Te can dissociate from the alloy and form Te_3

chains on the substrate in the continuous Te flux. And the Te_3 chains will self-assemble into a hexagonal lattice.⁸⁸

3.2.3 Wet-chemical synthesis. Wet-chemical synthesis, including the hydrothermal approach, has received less attention in the fabrication of these four Xene materials compared to epitaxial growth and exfoliation methods.

In 2019, few-layer β -antimonene with a hexagonal structure was successfully synthesized using SbCl_3 solutions as the source and oleylamine as the reducing agent, promoting their anisotropic growth in a colloidal solution (Fig. 11(a)). EDS reveals a small amount of S and O in the prepared gray Sb source from the precursor and surface oxidation, respectively. Dodecylthiol (DDT) and halide ions are identified as key contributors to avoiding the precursor transfer to Sb_2O_3 and to the formation of a few-layer antimonene, respectively (Fig. 11(a)).³⁴ The air-stable 2D tellurene, ranging from a monolayer to tens of nanometres, is successfully prepared *via* a hydrothermal method conducted in a high-pressure autoclave, with PVP and Na_2TeO_3 as precursors, and deionized water as the solvent. Controlling PVP concentration is the key to obtaining 2D tellurene, and the growth outcomes for different PVP concentrations

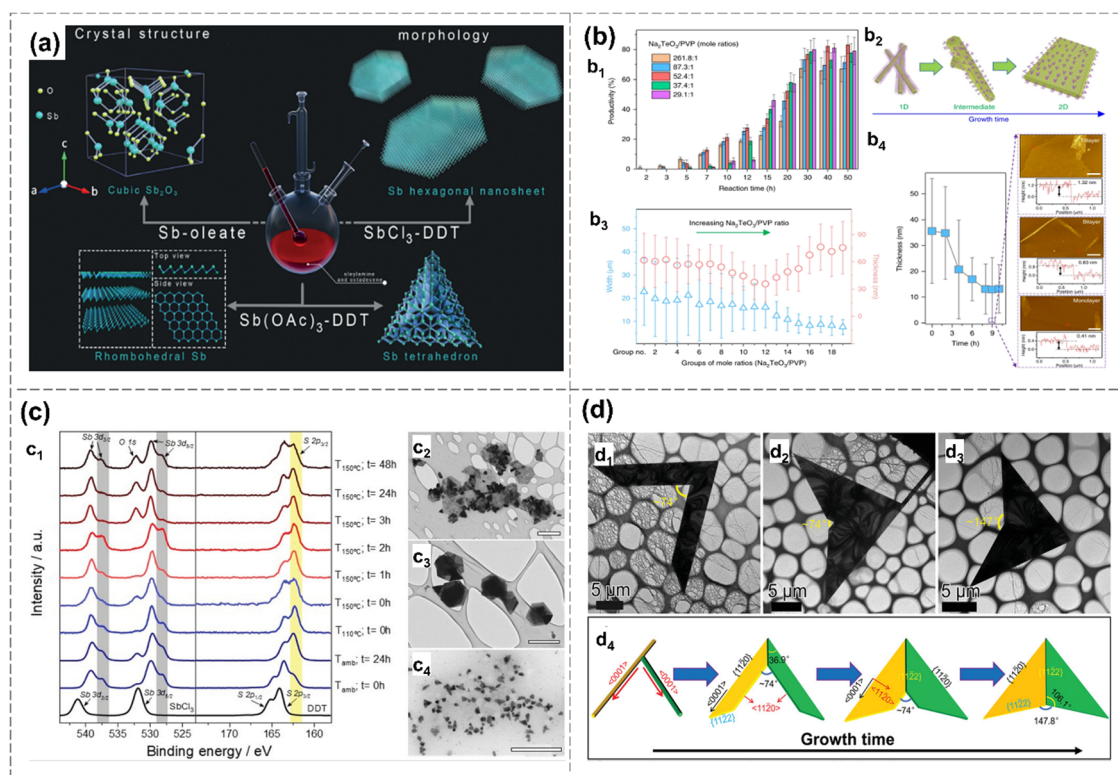
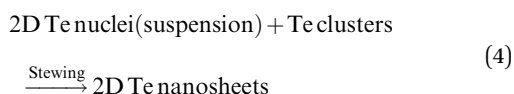
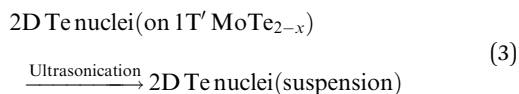
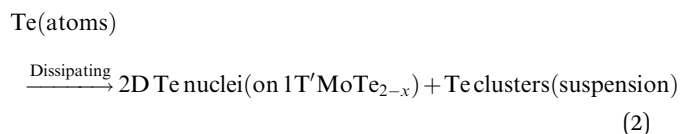


Fig. 11 Wet-chemical synthesis of antimonene and tellurene. (a) Schematic diagram of the wet-chemical synthesis of hexagonal antimonene, antimony trioxide and antimony tetrahedral.³⁴ (b) Solution growth and process of tellurene.¹²⁰ Productivity of tellurene over reaction time with different PVP concentrations (b₁). Morphology evolution from Te nanowires to 2D tellurene over growth time (b₂). Thickness and width of Te over mole ratios ($\text{Na}_2\text{TeO}_3/\text{PVP}$) (b₃). The thickness of 2D tellurene over a reaction time of thinning (b₄). (c) Characterization of 2D antimonene prepared by colloidal approach.¹⁰⁶ XPS spectra of Sb and S under different reaction conditions (c₁). TEM image of antimonene prepared at ambient temperature and t = 0 h (c₂). Sb particles obtained after t = 2 h (c₃) and t = 48 h (c₄).¹⁰⁶ (d) Synthesis mechanism of symmetric Te nanoflakes.¹²¹ TEM images of the "V-shaped" Te nanoflake (d₁), "heart-shaped" Te nanoflake (d₂) and "paper airplane-shaped" Te nanoflake (d₃). Illustration of morphology evolution from Te nanowire to symmetrical Te nanoflake (d₄). Reproduced from: (a) ref. 34 with permission from Wiley-VCH, Copyright 2019; (b) ref. 120 with permission from Springer Nature, Copyright 2018; (c) ref. 106 with permission from Wiley-VCH, Copyright 2021; (d) ref. 121 with permission from Wiley-VCH, Copyright 2021.

with varying times of reaction are shown in Fig. 11(b). The synthesized Te nanoflakes display remarkable optoelectronic performance, including high hole mobilities up to $458 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ under ambient conditions, exceptional anisotropic behavior and a broad bandwidth under $1.55 \mu\text{m}$ communication wavelength.¹²⁰ Another study achieves the synthesis of Te nanoflakes with a trigonal phase *via* hydrothermal methods. Due to shared identical growth directions between two nanoflakes, the nearly symmetrical morphologies are formed, including “V-shapes”, “paper airplane-shapes” and “heart-shapes” (Fig. 11(d)).¹²¹ The growth mechanism of 2D Te nanosheets is explained based on experimental evidence. Intersecting nanowires grow along respective $\langle 0001 \rangle$ direction initially, and then the growth direction is well-controlled using the passivation effect of PVP, leading to a transition from 1D nanowires to quasi-2D symmetrical nanoflakes.¹²¹



Notably, Qi *et al.* developed a novel ultrasonication-assisted synthesis method for the successful preparation of tellurene.⁹ As shown in Fig. 12, $1\text{T}'$ -phase MoTe_2 is applied as a Te source as well as the 2D growth template of 2D tellurene, eliminating the requirement for ligands or surfactants and ensuring environmental friendliness. The proposed synthesis mechanism is depicted in formula (1)–(4). $1\text{T}'$ - MoTe_2 initially undergoes an ultrasonication treatment in NMP solvent, which facilitates the delamination of van der Waals crystal. Then, a mild ultrasonication process is applied to disperse the mixture and generate the stable suspension of 2D tellurium nanosheets. The synthesis mechanism involves multiple interactions between energy and materials, encompassing crystal exfoliation, atomic escaping from the 2D lattice, preferential nucleation on 2D templates, and guided crystal growth. The resultant tellurene nanosheets exhibit single crystallinity with well-defined morphology and clean surfaces. 2D tellurium-based transistors demonstrate outstanding electrical performance, characterized by exceptional hole mobilities exceeding $1000 \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, and remarkable stability in both solid-state and liquid configurations.

4. Application of Xenes in catalysis

Antimonene has emerged as one of the most extensively studied Xenes for catalytic applications, closely followed by tellurene. Although pure Xenes generally possess limited catalytic activity due to their electron-enriched structure, which hinders the effective splitting of the molecule,¹⁶ Xenes can serve as promising substrates for atom loading to achieve enhanced catalytic performance. Furthermore, various engineering strategies, including the construction of van der Waals heterojunctions, the introduction of vacancies, interlayer modification, *etc.*, are employed to amplify the catalytic potential of Xenes.

Besides experimental studies, an increasing number of theoretical calculations are being used to predict the catalytic performance of Xenes and their composite materials. For example, a learning classifier utilizing synthetic minority over-sampling technology (SMOTE) is developed to identify 140 modified materials with promising HER properties from a pool of 25 low-cost transition metal and 2 nonmetal heteroatom dopants, all incorporated into gray arsenene systems. This selection is based on the principle of Gibbs free energy changes (DGH), which are less than 0.2 eV .⁶⁰ Subsequently, 13 types of Fe-single-atom-doped arsenene were developed as cost-effective catalysts following the principle of simple laboratory synthesis. This advanced machine learning approach addresses the inefficiencies of traditional high-throughput calculations and has demonstrated remarkable accuracy in predicting outcomes within complex systems. This work holds great promise for reducing labor costs and achieving efficient screening of the Xene catalyst systems. In this section, the discussion on the catalytic application of Xenes will be divided into electrocatalysis and photocatalysis.

4.1 Electrocatalysis

Wu *et al.* investigated the electrostatic gating, electrocatalytic properties, and electrochemical stability of tellurene across various electrolyte systems.³⁹ Tellurene synthesized *via* hydrothermal (h-Te) and chemical vapor deposition (c-Te) methods exhibits layer-dependent ambipolar transport properties, with notable high hole mobility and stability in both aqueous electrolytes and ionic liquids. Furthermore, the electrochemical performance of tellurene is investigated using the emerging on-chip electrocatalytic microdevices (OCEMs), which enable the detection of catalytic performance at the level of single nanosheets (Fig. 13(a)). In this study, AgCl/Ag serves as the reference electrode (RE), and the electrochemical potential is referenced to the reversible hydrogen electrode (RHE), as described by the following equation:

$$E_{\text{RHE}} = E_{\text{Ag}/\text{AgCl}} + 0.219 \text{ V} \quad (5)$$

E_{RHE} represents the potential of the reversible hydrogen electrode (RHE), and $E_{\text{Ag}/\text{AgCl}}$ represents the potential of the Ag/AgCl reference electrode.

Pristine α -phase tellurene is found to demonstrate excellent HER activity in acidic electrolytes, as indicated by its low Tafel



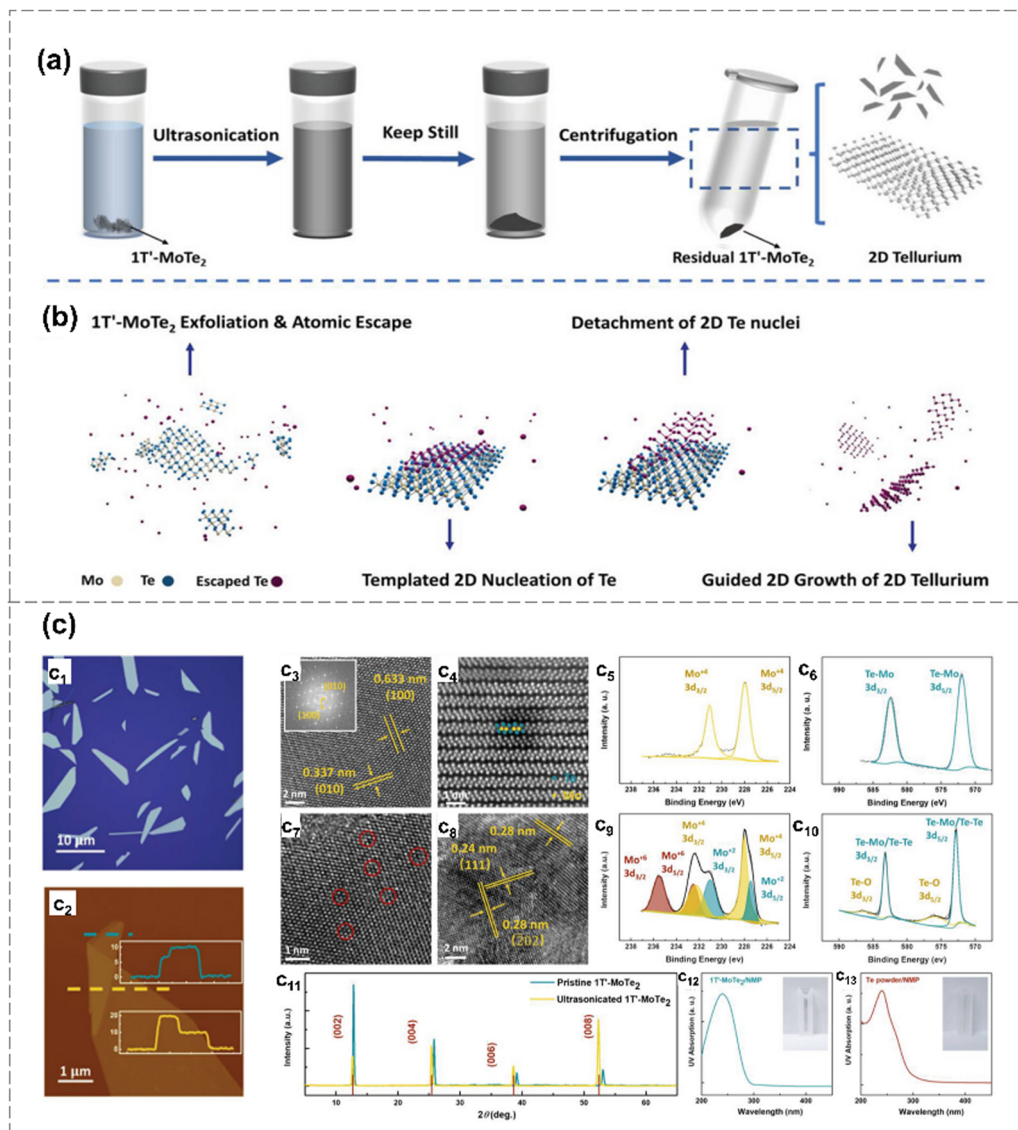


Fig. 12 Template growth of tellurene.⁹ (a), (b) Schematic diagram of tellurene synthesis using metastable 1T' MoTe₂ as the growth template (a) and corresponding proposed mechanism (b). (c) Characterization and growth mechanism of 2D Te. Optical microscopy of synthesized 2D tellurene (c₁). AFM image of a 2D tellurene with a thickness of 10 nm (c₂). HAADF-STEM images of 1T' MoTe₂ along (103) (c₃) and (100) (c₄) axis. XPS spectra of Mo 3d (c₅) and Te 3d (c₆) of 1T' MoTe₂. HRTEM image of 1T' MoTe₂ after ultrasonication. Red circles indicate the existence of defects (c₇). HRTEM image of polycrystalline MoO₃ on the surface of 1T' MoTe₂ (c₈). XPS spectra of Mo 3d (c₉) and Te 3d (c₁₀) of exfoliated 1T' MoTe₂. XRD pattern of pristine and exfoliated 1T' MoTe₂. UV-vis absorption spectra of the suspension 1T' MoTe₂/NMP (c₁₂) and bulk Te/NMP (c₁₃). Reproduced from ref. 9 with permission from Wiley-VCH, Copyright 2023.

slopes and high electrochemical stability across a wide potential window (Fig. 13). Such remarkable performance is attributed to the presence of vacancies, which act as active sites. Specifically, α -phase tellurene synthesized *via* the hydrothermal method exhibits an overpotential of 435 mV and a Tafel slope of 53 mV dec⁻¹. Moreover, few-layer antimonene is obtained *via* liquid exfoliation and employed as a non-metal bi-functional electrocatalyst for whole water splitting. It exhibits good catalytic activity and long-term stability in both OER and HER in an alkaline KOH solution.¹²² The lowest Tafel slope of few-layer Sb nanosheets is 261.3 mV dec⁻¹ for OER and 217.2 mV dec⁻¹ for HER.¹²²

Furthermore, in 2019, Lloret *et al.* reported the application of pristine Xenes in catalyzing organic reactions.¹²³ Few-layer, high-crystalline 2D Sb and BP with minimal defects in the zero-oxidation state are prepared by LPE using 1-butyl-3-methylimidazolium tetrafluoroborate (bmim-BF₄). These materials act as efficient catalysts for the alkylation of nucleophiles (such as alcohols, thiols and indoles) with simple esters under mild chemical environments. The use of Bmim-BF₄ and a glove box prevents oxidation and moisture-related degradation during the fabrication process, thereby ensuring the stability of antimonene. This research offers a more accessible alternative mechanism to the harsh reaction conditions

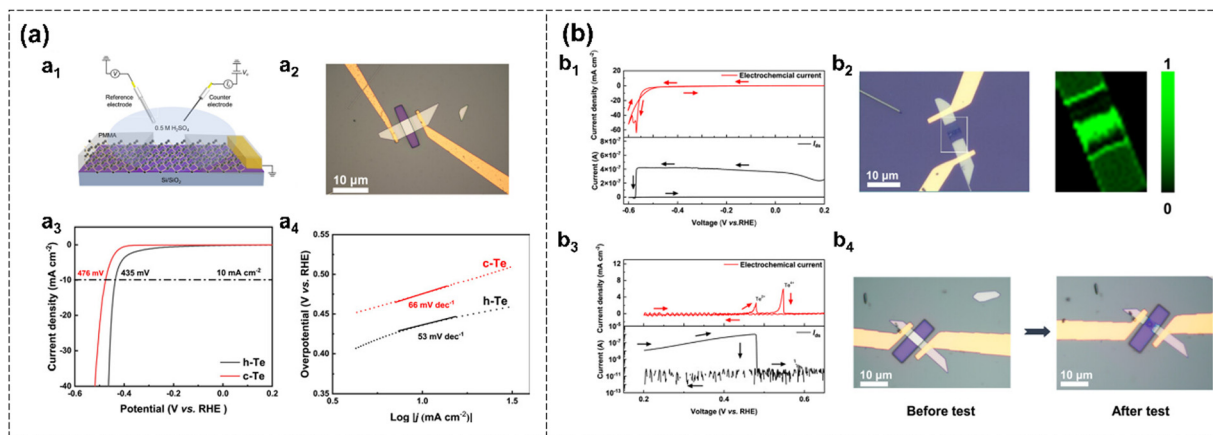


Fig. 13 (a) Electrocatalytic HER of α -phase tellurene investigated via OCEMs.³⁹ Schematic diagram of an OCEM for *in situ* transport measurements of tellurene (a₁). Optical microscopy image of h-Te devices with exposure area for testing (a₂). Polarization curves of h-Te (black) and c-Te (red) (a₃). Tafel plots of h-Te (black) and c-Te (red) obtained from a₃ (a₄). (b) Electrochemical stability of h-Te HER polarization curve and *in situ* transport property of h-Te in 0.5 mol L⁻¹ H₂SO₄ at V_{ds} = 100 mV (b₁). Optical image and Raman mapping of the h-Te device after the chemical test (b₂). Electrochemical curve (red) and transport curve (black) in the oxidation zone (b₃). Optical images of the h-Te device before and after the test (b₄). Reproduced from ref. 39 with permission from Wiley, Copyright 2023.

typically associated with superacid-catalyzed alkylation with esters. Notably, 2D antimonene outperforms BP due to its higher polarization effect, which enables the selective alkylation of acid-sensitive aromatic derivatives with simple esters. Additionally, a theoretical study suggests that the β -arsenene monolayer could exhibit an ultralow thermodynamic overpotential of 0.068 V for the anodic chlorine evolution reaction (CER).¹²⁴ The formation of antimony oxides on the surface can enhance its catalytic performance. To be specific, 2D Sb nanoflakes are used as an efficient electrocatalyst for the nitrogen reduction reaction (NRR), with a high NH₃ production rate of 180.4 $\mu\text{g h}^{-1} \text{mg}_{\text{CAT}}^{-1}$ and 11.6% faradaic efficiency (FE) in 0.1 M KOH solution under ambient environments. The theoretical calculations indicate that the oxidized species of antimonene serve as the active catalytic sites during the NRR process, and the increasing polarity resulting from facial oxidation makes the absorption energy of N₂ more favorable.⁴⁵

Additionally, antimonene is applied as a cocatalyst substrate and loaded with Pt atoms through the intercalation and reduction facilitated by ethylenediamine (EDA).¹¹⁴ The presence of Sb as the cocatalyst alters the oxidation pathway of formic acid on the Pt catalyst, favoring the dehydrogenation pathway reaction.¹¹⁴ A computational study indicates that Pt supported on an Sb substrate can be effectively activated to lower oxidative states, promoting the electrooxidation of HCOOH and reducing the CO poisoning effect. The as-prepared Pt/Sb exhibits a 4.5-fold increase in specific activity for the formic acid oxidation reaction (FAOR), and 2.6-fold higher mass activity compared to the conventional Pt/C catalysts.¹¹⁴ Antimonene has also been employed as a support for RuS₂ nanoparticles, forming Ru-S-Sb bonds.⁷¹ The RuS₂ nanoparticles exhibit *in situ* growth on sonication-exfoliated Sb nanosheets, demonstrating more stable catalytic abilities than Pt/C even after 40 hours of operation. The Ru-S-Sb/Sb heterostructure achieves a faradaic efficiency of 100% and displays

low Tafel slopes and high current density in both acidic (0.5 M H₂SO₄) and alkaline (1.0 M KOH) electrolytes. DFT calculations reveal that the formation of Ru-S-S promotes charge density around Ru-H, facilitating hydrogen absorption and reducing the HER thermodynamic barrier on the RuS₂ surface.⁷¹

Doping and defect engineering are the two most widely reported strategies for optimizing the catalytic performance of arsenene, antimonene and tellurene. For example, various transition metal (TM) single atom-doped arsenene is theoretically evaluated, with V-, Fe-, Co- and Ru-doped arsenene showing promise as electrocatalysts for the nitrogen reduction reaction (NRR) due to high TM loadings and stable nitrogen adsorption.⁵⁹ Oxygen is the most promising dopant used in black and gray arsenene to improve their ORR and OER catalytic performance.³⁷ Pt-doped Sb monolayer¹²⁵ and Bi-doped β -Sb⁷⁴ demonstrate potential as trifunctional electrocatalysts for HER, OER and ORR. Furthermore, arsenene can act as an electron acceptor and donor, with the VAs₃ (V and its attached three arsenic atoms) species serving as a charge transmitter between the adsorbed N_xH_y species and arsenene nanosheet during the NRR process.⁵⁹

Tellurene is synthesized through cathodic exfoliation which simultaneously facilitates the doping of transition metals.²⁶ Both Pt-doped and Ni-doped tellurene exhibit improved HER activities at high overpotentials, surpassing the performance of polycrystalline Pt. This behavior is attributed to the stable nature of tellurene, which allows for isotropic electron transfer at high mobility, facilitated by the delocalization of lone-pair electrons as well as the formation of strong Te-metal interactions for site immobilization.²⁶ The proposed catalytic mechanism involves the reduction of Te⁰ species to Te₂²⁻ with increasing overpotential, followed by the absorption of H⁺ to form H⁺-Te₂²⁻. The H⁺-Te-Pt-H_{ad} local structure facilitates the Heyrovsky step. After H₂ is released from the Pt sites, the surrounding H⁺-Te can provide protons for reduction on Pt.



Once Te_2^{2-} species are formed on the surface, the local adsorption–transfer–reduction–release processes are conducted collaboratively by the Te–Pt unit, enhancing the activity of individual metal active sites.²⁶

Additionally, Cao *et al.* found that defects in antimonene are introduced through LPE,⁷² with the density of defects increasing with prolonged treatment time. The resulting antimonene exhibits a high NH_3 yield ($2.08 \mu\text{g h}^{-1} \text{cm}^{-2}$) and faradaic efficiency (14.25% at -0.7 V vs. RHE). DFT calculations suggest that sufficient exposure to edge defects reduces the reaction barrier and enhances the interaction between antimonene and NRR intermediates, thereby improving ammonia selectivity and yield rate.⁷²

Moreover, Mishra *et al.* applied mechanical vibration to enhance the HER performance of Te. The refined Te nanosheets prepared *via* LPE exhibit an ultrahigh H_2 production rate of approximately $9000 \mu\text{mol g}^{-1} \text{h}^{-1}$ under ultrasonication due to its built-in electric field.⁹⁹ The rate of hydrogen production exhibits a positive correlation with the ultrasonication power and frequency. The measured piezocurrent ($\approx 2.5 \mu\text{A cm}^{-2}$) is significantly comparable to well-established piezoelectric materials for hydrogen generation so far.⁹⁹ The remarkable piezocatalytic performance of Te nanosheets arises from a synergistic interplay between their semi-metallic nature, favorable free energy landscape, enhanced electrical conductivity and outstanding piezoelectricity. As shown in Fig. 14(b), the estimated conduction band minimum (CBM) of Te nanosheets at -0.473 eV lies significantly below the hydrogen reduction potential (0 eV). And DFT results indicate that the gradual exposure of active sites on the Te nanosheets leads to a self-optimized catalytic performance for hydrogen generation.

Therefore, mechanically driven Te emerges as a novel piezo-catalyst for futuristic energy applications.

4.2 Photocatalysis

The large surface area and thickness-dependent bandgap of Xenes enable it with the potential to serve as a highly efficient photocatalyst.^{14,126} Both pristine Xene and its composites are reported to exhibit promising catalytic properties. Engineering strategies, such as doping and constructing heterostructures, are commonly employed to obtain Xenes with better photocatalytic performance.

For instance, pristine antimonene nanosheets are predicted to demonstrate photocatalytic properties across the entire pH range,¹²⁷ while Bi-doped β -antimonene monolayer is predicted to be a highly efficient trifunctional photocatalyst for ORR and overall water splitting.⁷⁴ Defect engineering of 2D oxidized Sb nanosheets is utilized to enhance the visible light-driven photocatalytic reaction for NRR (reduce N_2 to NH_3) under ambient pressure and temperature by introducing exposed active sites (Sb and O vacancies). The yield of NH_3 could be finely tuned by modifying the location and density of defects as well as changing the thickness of antimonene (Fig. 15). Remarkably, the NH_3 yield rate of the defective 2D photocatalyst is $388.5 \mu\text{g h}^{-1} \text{mg}_{\text{CAT}}^{-1}$ under visible light illumination without cocatalysts, significantly surpassing that of bulk Sb and Sb_2O_3 .¹²⁸ The ultrahigh activity of Sb is attributed to the extensive exposure of surface catalytic active sites associated with atomic thickness and defects, which promotes the migration and separation of charge carriers.

Furthermore, tellurene, antimonene and arsenene can be combined with other materials to enhance catalytic

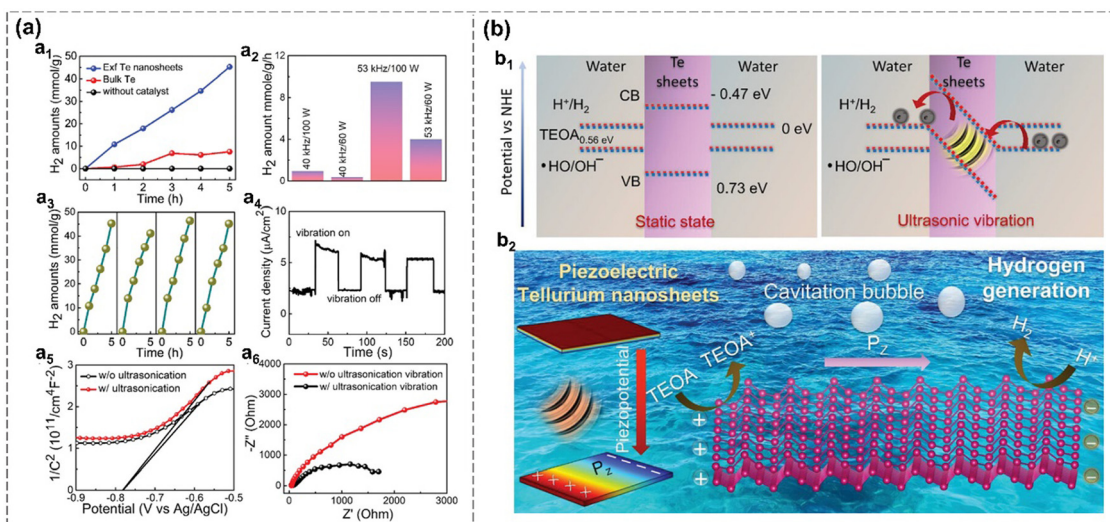


Fig. 14 Piezocatalytic HER in Te nanosheets.⁹⁹ (a) The electrochemical properties of Te nanosheets. H_2 production amount from bulk Te, refined exfoliated Te nanosheets, and blank control group under ultrasonic vibration over time (a_1). H_2 production rate over different mechanical vibration frequencies (a_2). Cyclic repeatability of H_2 production (a_3). Piezocurrent density over time with on and off vibration (a_4). Capacitance (a_5) and EIS spectra (a_6) with and without mechanical vibration. (b) Mechanism of piezocatalytic hydrogen generation in Te nanosheets. Energy band diagrams illustrate the charge transfer under static and ultrasonic vibration (b_1). Schematic diagram of the mechanism of piezocatalytic reaction for hydrogen production in Te (b_2). Reproduced from ref. 99 with permission from Wiley-VCH, Copyright 2024.



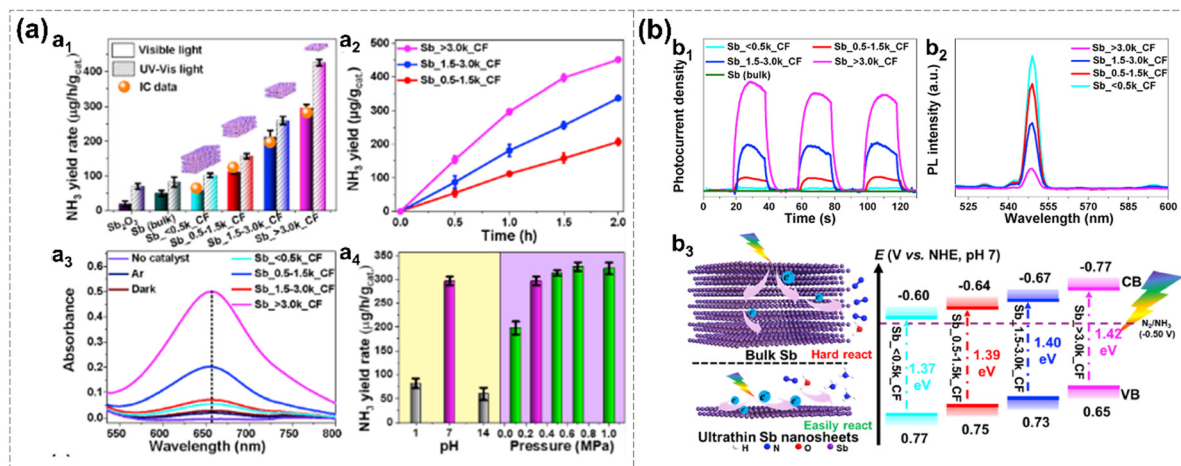


Fig. 15 Photocatalytic N_2 fixation and the energy band structures of Sb nanosheets.¹²⁸ NH_3 yield rates of Sb nanosheets over different cascade centrifugation (CF) speed tested by the indophenol blue method and ion chromatography (orange balls) (a₁). NH_3 yield of Sb nanosheets over reaction time (a₂). Absorption spectra of Sb nanosheets in the dark, Ar atmosphere (under visible light), and without catalysts (under visible light) (a₃). NH_3 yield rate of Sb_{>3.0k}_CF under different pH values and pressure (a₄). The photocurrent density of Sb nanosheets under different CF over time (b₁). PL intensity of Sb nanosheets under different CF (b₂). Schematic diagram of N_2 fixation and the energy band structure of Sb nanosheets (b₃). Reproduced from ref. 128 with permission from Elsevier, Copyright 2020.

performance. They can act as the substrates for single atom catalysts (SACs) to interact with single-atom active sites or be used to form heterostructures with other layered 2D materials to achieve synergistic effects. For example, tellurene nanosheets are employed as metal-free cocatalysts and combined with CdS through a one-step hydrothermal method.¹²⁶ The number of incident photons and quantum efficiency (QE) are calculated based on the following eqn (6) and (7), respectively.

$$N = \frac{E\lambda}{hc} \quad (6)$$

$$QE = \frac{2 \times \text{the number of the evolved } H_2 \text{ molecules}}{\text{The number of incident photons}} \quad (7)$$

This photocatalytic system exhibits a hydrogen production rate of $234.4 \text{ mmol h}^{-1}$, a 7.8-fold improvement in catalytic activity compared to bare CdS (30.2 mmol h^{-1}). A high wavelength-dependent apparent QE (AQE) of 11.4% and 5.6%, can be obtained at the fixed wavelength of 420 nm and 475 nm in prepared Te/CdS, respectively, which is comparable to most of the CdS merged with the previously reported metal-free and noble-metal-free cocatalysts.¹²⁶ The formed favorable band alignment with photocatalysts and its thickness-dependent photocatalytic performance is investigated, revealing efficient electron-hole separation in the Te/CdS system.¹²⁶ This is attributed to the fast electron transfer ability of tellurene suppresses the recombination of photo-generated charge on CdS.¹²⁶

Additionally, atomic cobalt species (ACS) anchored to 2D tellurene are demonstrated as efficient photocatalysts for both the HER and CO_2 reduction reactions (CO_2RR). Tellurene acts as an active single-atom cocatalyst, anchoring ACS and bridging between the light absorber and ACS active sites for

efficient electron transfer.²⁷ The Co-Te bond forms between ACS and Te substrate, and its strong interaction leads to altered electronic structures in Te, creating intermediate energy states that can catch the photogenerated electron for better photocatalytic performance. As a result, the Te NS-ACS exhibit an outstanding enhanced visible-light photocatalytic H_2 production rate of $130.8 \mu\text{mol h}^{-1}$ in the presence of Eosin Y (EY) as a light absorber, and a CO production rate of $52.3 \mu\text{mol h}^{-1}$ in the presence of tris(2,2'-bipyridyl) dichlororuthenium(II) ($[Ru(bpy)_3]Cl_2$) as a light absorber, representing the state-of-the-art noble-metal-free cocatalyst²⁷ (Fig. 16).

Furthermore, 2D/2D graphitic carbon nitride/antimonene type I heterostructure (CNSb_x) is prepared *via* a two-step procedure. It shows promoted light absorption, electron-hole separation capabilities, and higher photocatalytic activity for the degradation of organic pollutants, including rhodamine B (RhB) and *p*-nitrophenol (*p*-NP), compared with pristine carbon nitride (CN).¹²⁹ The photochemical *N*-deethylation of RhB occurs through oxidation facilitated by photogenerated holes from the valence band of CN. However, electrons transferred from the conduction band of graphitic carbon nitride (g-CN) to the conduction band of Sb are also involved during the degradation process.¹²⁹

The arsenene/g-C₃N₄ van der Waals heterostructure is theoretically studied for its potential application in water splitting.¹³⁰ It exhibits improved optical absorption in both the visible light and ultraviolet regions compared to pristine monolayers. Thanks to their type-II band alignment and low exciton binding energy, they feature a solar-to-hydrogen (STH) efficiency limit of up to 15.8%, facilitating the effective separation of photogenerated electrons and holes.¹³⁰ Besides, PtS₂/arsenene van der Waals heterostructure is demonstrated to exhibit a direct Z-scheme catalytic.⁵⁷ This PtS₂/arsenene heterostructure is predicted to achieve an enhanced solar-to-hydrogen



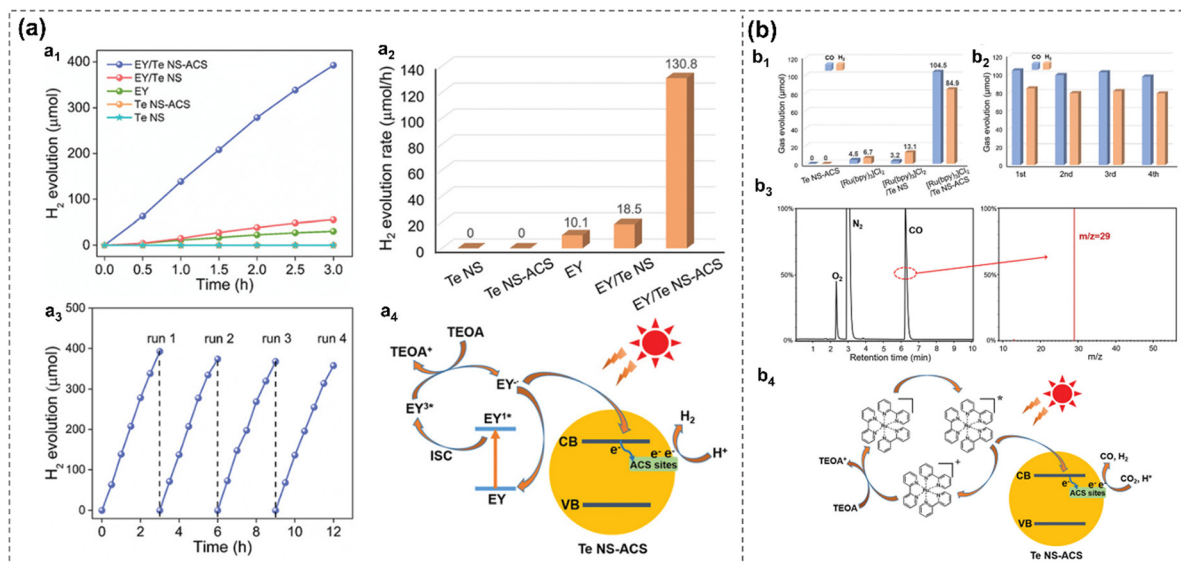


Fig. 16 Photocatalytic performance and proposed mechanism of EY/Te NS-ACS (b) and [Ru(bpy)₃]Cl₂/Te NS-ACS (a), respectively.²⁷ Time-dependent H₂ evolution (a₁) and H₂ evolution rate (a₂) of different materials. Cycling test of EY/Te NS-ACS for H₂ evolution (a₃). Proposed mechanism of EY/Te NS-ACS for H₂ evolution (a₄). CO and H₂ evolution from CO₂RR of different materials (b₁). Cycling test of [Ru(bpy)₃]Cl₂/Te NS-ACS from CO₂RR (b₂). GC-MS (JEOL-GCQMS, JMS-K9, and 6890N Network GC) result of ¹³C generated in the reduction of ¹³CO₂ over [Ru(bpy)₃]Cl₂/Te NS-ACS (b₃). Proposed mechanism for photocatalytic CO₂RR over [Ru(bpy)₃]Cl₂/Te NS-ACS (b₄). Reproduced from ref. 27 with permission from Wiley-VCH, Copyright 2020.

(STH) efficiency (49.32%) compared with isolated PtS₂ (12.67%) or arsenene (10.34%) monolayers.⁵⁷

5. Summary and perspectives

Antimonene, arsenene, tellurene and selenene possess remarkable inherent properties, including large specific areas, numerous active sites, outstanding conductivity, tunable bandgaps, and robust stability compared to phosphorene, making them highly promising materials for various catalytic applications. This review focuses on the recent advancements (publications from 2019 to 2023) concerning these four emerging Xenes. In addition to introducing the stable phases of these Xenes, along with their layer-dependent properties, particularly with regard to bandgap tunability, diverse preparation techniques, including top-down methods and bottom-up methods, are discussed. Engineering methods such as heterostructures, doping, defects and external strain fields, and applications in electrocatalysis, photocatalysis including HER, OER, ORR, CO₂RR, NRR, and organic catalysis are introduced.

It is noteworthy that these four Xenes are generally believed to exhibit limited inherent catalytic activity. Specifically, black and gray arsenene exhibit low catalytic performance due to their reduced molecular absorption capabilities. Additionally, these materials exhibit suboptimal photocatalytic ability owing to their unsuitable bandgaps, challenges in electron-hole separation, and rapid recombination rates. While the bulk form of these Xenes is reported to be stable, their monolayer or bilayer counterparts are prone to degradation in ambient environments. This instability complicates the acquisition of pure, defect-free forms of these four Xenes and poses

challenges to the evaluation of their inherent physical and chemical properties. Moreover, reports on the large-scale synthesis of these Xenos are also limited, which hinders their industrial catalytic applications.

To address these challenges, scientists develop various engineering approaches to enhance the performance of these Xenos. For example, building heterostructures can minimize electron-hole recombination and align band edges with the requisite redox potentials for photocatalysis. Additionally, utilizing the large surface area, abundant active sites and high carrier mobility of antimonene to support transition metal SACs can improve gas molecular absorption and lower the overpotential. Moreover, machine learning and other advanced techniques are employed to identify optimal structures and properties for heterostructures and element doping combinations, streamlining the process and reducing manual labor. Furthermore, external safety measures such as protective coating are employed to minimize degradation during fabrication and storage.

Despite significant contributions, achieving large-scale synthesis of these Xenos with extended lateral dimensions remains a challenge. Additionally, a comprehensive understanding of the complex relationship between processing methods, structural changes, and resulting properties is crucial for bridging the gap between theoretical predictions and practical applications.

Author contributions

Cong Ma: conceptualization, methodology, formal analysis, investigation, writing – original draft; Kai Bao: formal analysis,

validation, investigation, writing – review & editing; Junlei Qi: validation, writing– review & editing, supervision; Wenbin Wang: validation, investigation, resources; Jingkun Wu: validation, investigation; Ruijie Li: writing – review & editing, supervision; Yu Lin: writing – review & editing, supervision; Lingzhi Wang: validation; Zhuangzhuang Yin: validation; Zhixiang Tao: validation; Chengxuan Ke: supervision; Zongxiao Wu: supervision; Yingxia Liu: writing – review & editing, supervision, funding acquisition; Qiyuan He: writing – review & editing, supervision, funding acquisition.

Data availability

No primary research results, software or code have been included and no new data were generated or analysed as part of this review.

Conflicts of interest

The authors declare that they have no known competing financial interests or personal relationships that could have appeared to influence the work reported in this paper.

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