

Cite this: *Nanoscale*, 2019, **11**, 19297Received 22nd July 2019,
Accepted 9th September 2019

DOI: 10.1039/c9nr06239d

rsc.li/nanoscale

A facile synthetic approach to nanostructured Li₂S cathodes for rechargeable solid-state Li–S batteries†

Hany El-Shinawi,^a Edmund J. Cussen^b and Serena A. Corr^{a*}

Li–S solid state batteries, employing Li₂S as a pre-lithiated cathode, present a promising low cost, high capacity and safer alternative to their liquid electrolyte counterparts, where dissolution of intermediate polysulfide species can result in loss of active material and a subsequent decrease in ionic conductivity. A nanostructured Li₂S material would afford greater flexibility in optimising the cathode composite for more harmonious electrode–electrolyte interactions, yet facile routes to such nanoscale materials are limited. Here, we report a facile and scalable microwave approach to directly synthesize nanostructured Li₂S from a glyme solution containing lithium polysulfides. As-synthesized Li₂S presents an ideal architecture for the construction of free-standing cathodes for all-solid-state Li–S batteries.

Reversible electrochemical lithiation of S to Li₂S in rechargeable Li–S batteries is a promising approach to realizing low-cost, high energy density batteries.¹ Abundant, light-weight sulfur offers a potential specific capacity of 1672 mA h g^{−1} which, at an order of magnitude higher than that provided by conventional intercalation-type cathodes such as LiCoO₂ and LiFePO₄, makes Li–S batteries suitable for high energy storage applications such as powering electric vehicles.^{1,2} Li₂S has been suggested as a pre-lithiated cathode, negating the necessity of metallic lithium as an anode.^{2,3} With the application of a solid-electrolyte, further advantages such as higher ionic conductivities and potentially mitigating deleterious polysulfide shuttle activity are possible.^{4–6} However, a serious bottleneck in the widespread uptake of such a pre-lithiated material has been the limited synthetic routes to Li₂S. Preferably, Li₂S would be composited with an electronic conductor and a Li-ion conducting medium on the nanoscale to promote greater

electrode–electrolyte interactions, in an architecture that would accommodate the volume changes associated with cycling. The application of nanostructured Li₂S also Recent years have seen efforts made in developing novel Li₂S composites, for example in combination with carbon and/or solid electrolyte.³ A significant benefit to the community would be a reliable and scalable synthetic route to nanostructured Li₂S, which up to now has been limited and in several cases complicated and costly. Reported routes include impregnating carbon materials with ethanolic solutions of commercial Li₂S,^{7–16} chemical reaction of sulfur with lithium triethylborohydride or stabilized lithium metal powder,^{4,17–23} carbon-based reduction of Li₂SO₄ and polysulfides,^{24–30} reactions between H₂S and lithium naphthalenide,³¹ and burning Li foils in a CS₂ vapour.³² Thus, there remains a lack of facile routes to Li₂S which are scalable and afford control over the final particle morphology and size. Here we report the synthesis of nanostructured Li₂S, prepared by a simple microwave approach, together with its performance in an all solid-state Li–S battery. Our results demonstrate that nanostructured Li₂S employed within a free-standing cathode displays exceptional cycling stability, achieving capacities of 440 mA h g^{−1} at cycling rates of 100 μA cm^{−2} up to 400 cycles.

Nanostructured Li₂S is prepared using a simple 20-minute microwave-assisted heat treatment of a tetraglyme solution containing elemental sulfur and lithium *tert*-butoxide, at temperatures as low as 200 °C (see ESI for experimental details†). X-ray diffraction patterns of this microwave-synthesised Li₂S powder (Fig. 1) reveal peak broadening consistent with a reduction in particle size (with an estimated crystallite size of 10 nm from Scherrer broadening), compared to a commercial Li₂S sample (Sigma Aldrich, 99.98%). All peaks could be indexed to the cubic lattice of the antifluorite structure of Li₂S. Nanoparticles of Li₂S are confirmed by scanning electron microscopy (SEM) images, where aggregates of particles are observed. Fig. 2 illustrates the importance of the S/lithium *tert*-butoxide concentration ratio on the resulting particle size and distribution, with three examples shown (see Fig. S1† for

^aDepartment of Chemical and Biological Engineering, University of Sheffield, Sir Robert Hadfield Building, Sheffield, S1 3JD, UK. E-mail: s.a.corr@sheffield.ac.uk

^bDepartment of Materials Science and Engineering, University of Sheffield, Sir Robert Hadfield Building, Sheffield, S1 3JD, UK

†Electronic supplementary information (ESI) available. See DOI: 10.1039/c9nr06239d



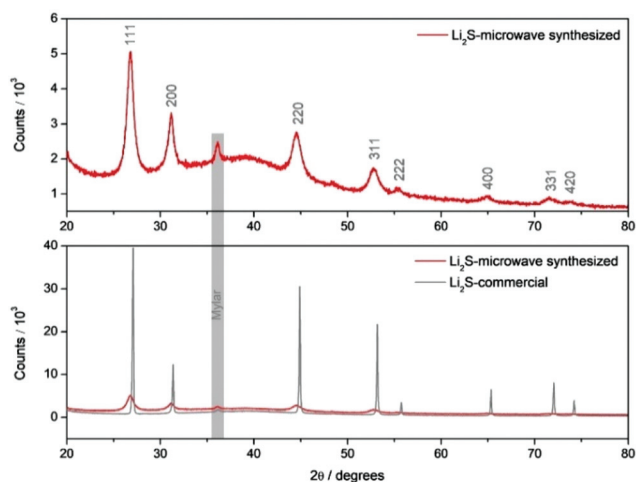


Fig. 1 XRD patterns of commercial Li_2S and microwave-synthesized Li_2S (0.15 M sulfur solution). An air tight sample holder employing a Mylar window was employed.

further SEM images). Our experiments reveal that Li_2S prepared using 0.15 M sulfur solution gave particles with the smallest size and narrowest size distribution.

Such nanostructured material provides a promising electrode architecture in terms of providing shorter pathlengths for Li-ion transport, as well as larger electrode–electrolyte contact areas, compared with bulk materials and this sample was selected for further electrochemical study.

It has been previously observed that conductive thin layers of Li_3PS_4 can be deposited from a THF solution containing Li_2S and P_2S_5 .^{4,33} Here, we followed a modified approach to incorporate our microwave-synthesized nanostructured Li_2S in a three-phase $\text{Li}_3\text{PS}_4/\text{nano-Li}_2\text{S}/\text{C}$ nanocomposite. Firstly, nanostructured carbon black is composited with P_2S_5 using a melt-diffusion technique (Fig. S2†) before mixing with microwave-synthesized Li_2S in THF (molar ratio of $\text{P}_2\text{S}_5:\text{Li}_2\text{S}$ is 1:20) to produce the final $\text{Li}_3\text{PS}_4/\text{Li}_2\text{S}/\text{C}$ nanocomposite. This corresponds to 40 wt% Li_2S in the composite. Characterisation of the final composite by SEM, EDX and XRD are shown in Fig. 3. EDX analysis reveals a homogeneous phosphorus and sulfur distribution across the material, while the presence of Li_3PS_4 is not evident from XRD nor from SEM (pure $\beta\text{-Li}_3\text{PS}_4$ has a distinctive morphology of micron-sized, porous brick-shaped particles). These data suggest the formation of a thin film of Li_3PS_4 , which is either amorphous or is formed on a much shorter length scale not detectable by Bragg diffraction. The formation of a Li_3PS_4 thin coating on the $\text{Li}_2\text{S}/\text{C}$ surface is

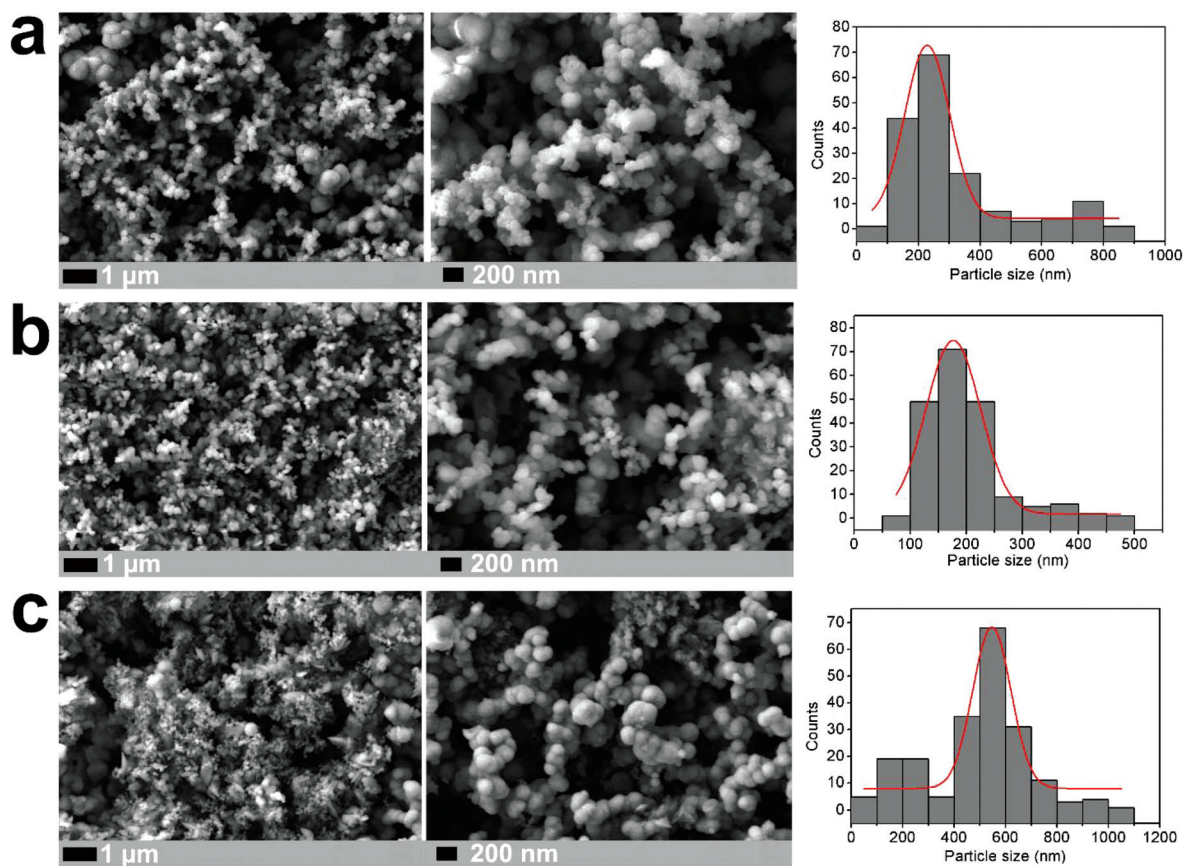


Fig. 2 Morphology and particle size distributions based on SEM ($N = 170$) of Li_2S prepared from (a) 0.075 M, (b) 0.15 M and (c) 0.3 M sulfur solutions. The mean particle sizes are given as $228 (\pm 72)$ nm, $177 (\pm 49)$ nm and $546 (\pm 74)$ nm respectively.



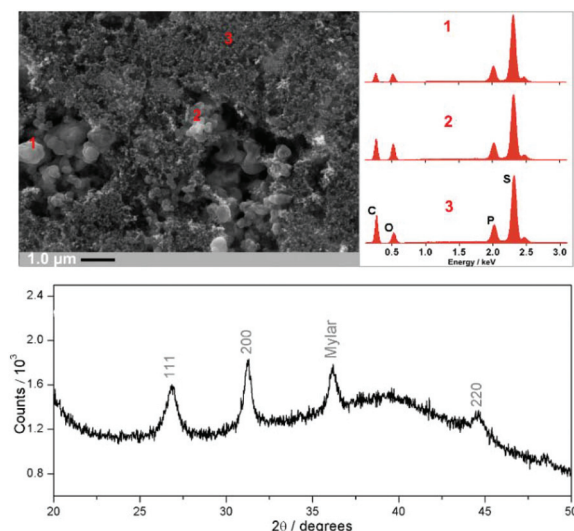


Fig. 3 SEM/EDX analysis and XRD pattern of the as-synthesized $\text{Li}_3\text{PS}_4/\text{Li}_2\text{S}/\text{C}$ composite.

consistent with previous reports.⁴ This is further confirmed by performing the control experiment, in which a P_2S_5 -rich/C composite is mixed with Li_2S (molar ratio of $\text{P}_2\text{S}_5:\text{Li}_2\text{S}$ is 1 : 10). In addition to formation of bulk $\beta\text{-Li}_3\text{PS}_4$, which is confirmed by XRD and SEM (Fig. S3†), a homogeneous distribution of P and S is also observed across this material indicating the formation again of a thin coating of Li_3PS_4 .

An all-solid-state battery was prepared using the $\text{Li}_3\text{PS}_4/\text{Li}_2\text{S}/\text{C}$ nanocomposite as a free-standing cathode, $\beta\text{-Li}_3\text{PS}_4$ as the solid electrolyte and metallic lithium as the anode. These components were pressed together under 2 tons of pressure, employing Al and Ni foils to support and collect the current from the cathode and the anode, respectively (Fig. 4a, insert). The cell cycled at 60 °C and at current densities ranging between 20 and 100 $\mu\text{A cm}^{-2}$ for up to 400 cycles (corresponding to $\sim C/2$ and $C/10$). Fig. 4a shows the typical charge and discharge profiles of the battery at 20 $\mu\text{A cm}^{-2}$. In addition to retaining the capacity, it is notable that the voltage profiles comprise well-defined discharge and charge plateaus, displaying small potential hysteresis. We attribute this small hysteresis to improved electronic and ionic conductivities of the composite cathode and a reduced interfacial resistance at the solid-electrolyte/cathode interface, as illustrated by impedance spectroscopy measurements (see below). Increasing the discharge/charge rate from 20 to 100 $\mu\text{A cm}^{-2}$ reduced the capacity from ~ 840 to ~ 440 mA h g^{-1} , but this capacity is retained over more than 300 additional cycles with coulombic efficiency close to 100% (Fig. 4b). Fig. S4† displays SEM images of the $\text{Li}_3\text{PS}_4/\text{Li}_2\text{S}/\text{C}$ electrode before and after cycling revealing no obvious change of the morphology after 410 cycles at different rates. Some surface cracks are observed which can be attributed to volume changes during cycling; importantly, these cracks do not lead to a loss of contact between the electrode domains or to delamination of the elec-

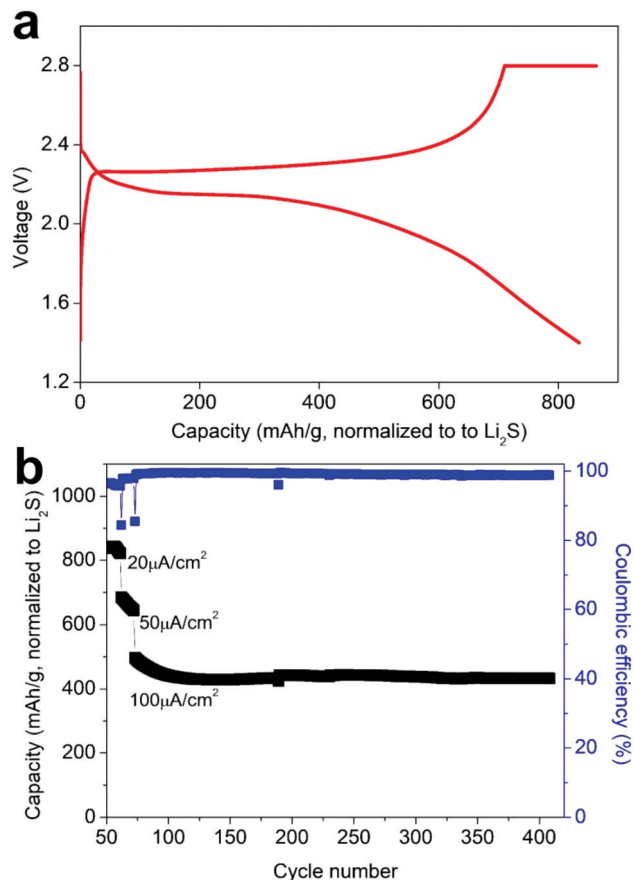


Fig. 4 (a) Typical charge/discharge profile of the all-solid-state battery at 20 $\mu\text{A cm}^{-2}$ current density. (b) Cycle performance of the battery at current densities of 20, 50 and 100 $\mu\text{A cm}^{-2}$ over 400 cycles.

trode and the solid electrolyte. Impedance measurements prior to and after cycling, indicated an increase of the resistance of the cell from $\sim 250 \Omega$ to $\sim 600 \Omega$ after 410 cycles at different rates (Fig. S5†). This limited growth of cell resistance with long-term cycling is consistent with capacity retention and maintaining electrode integrity and suggests good stability of the solid-electrolyte/cathode interface. For comparison, in a recent study on $\text{Li}_6\text{PS}_5\text{Cl}-\text{Li}_2\text{S}$ all-solid-state batteries, a cycled battery employing $\text{Li}_2\text{S}/\text{C}$ as the cathode exhibited an interfacial resistance that is one order of magnitude higher than the resistance of the fresh cell after only 37 cycles.³⁴

Conclusions

In summary, we demonstrate that nanostructured Li_2S can be prepared by a simple, fast, microwave heat treatment of a glyme solution containing lithium polysulfides. The as-synthesized Li_2S presents an ideal architecture for the construction of free-standing cathodes comprising $\text{Li}_3\text{PS}_4/\text{Li}_2\text{S}/\text{C}$ nanocomposites. Crucially, this synthetic approach is highly reproducible, having been repeated numerous times, and scalable. The performance of this nanostructured Li_2S as a cathode in



an all-solid-state battery demonstrates excellent capacity retention at current densities of up to $100 \mu\text{A cm}^{-2}$, with limited increases in cell resistance and close to 100% coulombic efficiencies, which we attribute to increased electrode–electrolyte contact area and shorter Li^+ diffusion pathlengths.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

We gratefully acknowledge the support of the University of Sheffield in our research. Our work was supported by the EPSRC [EP/N001982/2] and the ISCF Faraday Challenge project “SOLBAT – The Solid-State (Li or N (a) Metal-Anode Battery” [grant number FIRG007].

Notes and references

- (a) P. G. Bruce, S. A. Freunberger, L. J. Hardwick and J. M. Tarascon, *Nat. Mater.*, 2012, **11**, 19; (b) S.-K. Lee, Y. J. Lee and Y. K. Sun, *J. Power Sources*, 2016, **323**, 174.
- A. Manthiram, Y. Fu, S.-H. Chung, C. Zu and Y.-S. Su, *Chem. Rev.*, 2014, **114**, 11751.
- D. Su, D. Zhou, C. Wang and G. Wang, *Adv. Funct. Mater.*, 2018, **28**, 1800154.
- Z. Lin, Z. C. Liu, N. J. Dudney and C. D. Liang, *ACS Nano*, 2013, **7**, 2829.
- F. Han, J. Yue, X. Fan, T. Gao, C. Luo, Z. Ma, L. Suo and C. Wang, *Nano Lett.*, 2016, **16**, 4521.
- Z. Jiao, L. Chena, J. Si, C. Xu, Y. Jiang, Y. Zhu, Y. Yang and B. Zhao, *J. Power Sources*, 2017, **353**, 167.
- F. Wu, H. Kim, A. Magasinski, J. T. Lee, H.-T. Lin and G. Yushin, *Adv. Energy Mater.*, 2014, **4**, 1400196.
- F. Wu, A. Magasinski and G. Yushin, *J. Mater. Chem. A*, 2014, **2**, 6064.
- C. Wang, X. Wang, Y. Yang, A. Kushima, J. Chen, Y. Huang and J. Li, *Nano Lett.*, 2015, **15**, 1796.
- M. Wu, Y. Cui and Y. Fu, *ACS Appl. Mater. Interfaces*, 2015, **7**, 21479.
- F. Wu, J. T. Lee, F. Fan, N. Nitta, H. Kim, T. Zhu and G. Yushin, *Adv. Mater.*, 2015, **27**, 5579.
- Y. Qiu, G. Rong, J. Yang, G. Li, S. Ma, X. Wang, Z. Pan, Y. Hou, M. Liu, F. Ye, W. Li, Z. W. Seh, X. Tao, H. Yao, N. Liu, R. Zhang, G. Zhou, J. Wang, S. Fan, Y. Cui and Y. Zhang, *Adv. Energy Mater.*, 2015, **5**, 1501369.
- G. Zhou, E. Paek, G. S. Hwang and A. Manthiram, *Adv. Energy Mater.*, 2016, **6**, 1501355.
- J. He, Y. Chen, W. Lv, K. Wen, C. Xu, W. Zhang, W. Qin and W. He, *ACS Energy Lett.*, 2016, **1**, 820.
- F. Wu, J. T. Lee, E. Zhao, B. Zhang and G. Yushin, *ACS Nano*, 2016, **10**, 1333.
- D. H. Wang, D. Xie, T. Yang, Y. Zhong, X. L. Wang, X. H. Xia, C. D. Gu and J. P. Tu, *J. Power Sources*, 2016, **313**, 233.
- J. A. Gladysz, V. K. Wong and B. S. Jick, *Tetrahedron*, 1979, **35**, 2329.
- Z. Lin, C. Nan, Y. Ye, J. Guo, J. Zhu and E. J. Cairns, *Nano Energy*, 2014, **9**, 408.
- K. Zhang, L. Wang, Z. Hu, F. Cheng and J. Chen, *Sci. Rep.*, 2014, **4**, 6467.
- C. Nan, Z. Lin, H. Liao, M. K. Song, Y. Li and E. J. Cairns, *J. Am. Chem. Soc.*, 2014, **136**, 4659.
- Y. Hwa, J. Zhao and E. J. Cairns, *Nano Lett.*, 2015, **15**, 3479.
- L. Suo, Y. Zhu, F. Han, T. Gao, C. Luo, X. Fan, Y.-S. Hu and C. Wang, *Nano Energy*, 2015, **13**, 467.
- S. Zheng, Y. Chen, Y. Xu, F. Yi, Y. Zhu, Y. Liu, J. Yang and C. Wang, *ACS Nano*, 2013, **7**, 10995.
- Z. Yang, J. Guo, S. K. Das, Y. Yu, Z. Zhou, H. D. Abruña and L. A. Archer, *J. Mater. Chem. A*, 2013, **1**, 1433.
- Z. Li, S. Zhang, C. Zhang, K. Ueno, T. Yasuda, R. Tatara, K. Dokko and M. Watanabe, *Nanoscale*, 2015, **7**, 14385.
- J. Zhang, Y. Shi, Y. Ding, L. Peng, W. Zhang and G. Yu, *Adv. Energy Mater.*, 2017, **7**, 1602876.
- M. Yu, Z. Wang, Y. Wang, Y. Dong and J. Qiu, *Adv. Energy Mater.*, 2017, 1700018.
- F. Ye, H. Noh, J. Lee, H. Lee and H.-T. Kim, *J. Mater. Chem. A*, 2018, **6**, 6617.
- Y. Fu, C. Zu and A. Manthiram, *J. Am. Chem. Soc.*, 2013, **135**, 18044.
- M. Kohl, J. Brückner, I. Bauer, H. Althues and S. Kaskel, *J. Mater. Chem. A*, 2015, **3**, 16307.
- X. M. Li, C. A. Wolden, C. M. Ban and Y. G. Yang, *ACS Appl. Mater. Interfaces*, 2015, **7**, 28444.
- G. Tan, R. Xu, Z. Xing, Y. Yuan, J. Lu, J. Wen, C. Liu, L. Ma, C. Zhan and Q. Liu, *Nat. Energy*, 2017, **2**, 17090.
- H.-D. Lim, H.-K. Lim, X. Xing, B.-S. Lee, H. Liu, C. Coaty, H. Kim and P. Liu, *Adv. Mater. Interfaces*, 2018, **5**, 1701328.
- C. Yu, S. Ganapathy, N. J. J. de Klerk, I. Roslon, E. R. H. van Eck, A. P. M. Kentgens and M. Wagemaker, *J. Am. Chem. Soc.*, 2016, **138**, 11192.

