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Reversible disassembly of metallasupramolecular structures mediated by a metastable-state photoacid†

Suzanne M. Jansze, , Giacomo Cecot and Kay Severin *

The addition of a metastable-state photoacid to solutions containing metal–ligand assemblies renders the systems light responsive. Upon irradiation, proton transfer from the photoacid to the ligand is observed, resulting in disassembly of the metallasupramolecular structure. In the dark, the process is fully reversed. Light-induced switching was demonstrated for six different metal–ligand assemblies containing Pd^{II}, Pt^{II} or Ru^{II} complexes and bridging polypyridyl ligands. The methodology allows liberating guest molecules with light.

Introduction

The controlled manipulation of supramolecular structures with light represents an interesting option to create functional nanoscale devices.¹ A common strategy to implement light-responsiveness in metal–ligand assemblies is the utilization of photochromic ligands. Light allows altering the geometry of the ligand, which in turn may result in a structural reconfiguration of the metallasupramolecular assembly.^{1,2} For example, bipyridyl ligands containing dithienylethene units have been used for the light-induced conversion of a Pd₂₄L₄₈ capsule into small metallamacrocycles,³ to create Pd₂L₄ cages with light-dependent host–guest chemistry,⁴ and for the photo-induced sol–gel switching of a metallo-gel.⁵ Ligands containing the azobenzene motif were employed to modulate the host–guest chemistry of a spherical Pd₁₂L₂₄ complex,⁶ to alter the solubility

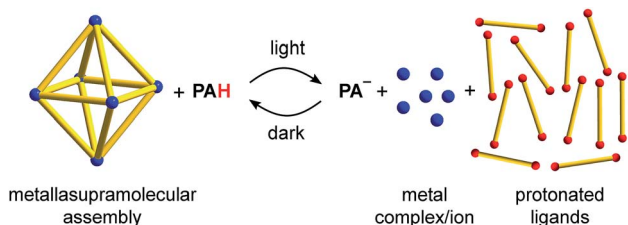
of a coordination cage,^{7,8} and to control the catenation of a palladium metallacycle.⁹

The construction of light-responsive metal–ligand assemblies with photochromic ligands is not straightforward. The design of suited ligands can be challenging, and the preparation of the ligands may involve substantial synthetic efforts. In this study, we describe an alternative and more general approach to render metal–ligand assemblies light-responsive: the addition of a metastable-state photoacid. We demonstrate that the acidification by the photoacid upon irradiation can be used for the disassembly of metallasupramolecular structures.¹⁰ The process is reversible, and the metallasupramolecular structures form again in the dark at room temperature (Scheme 1).

Results and discussion

For our investigations, we decided to use the merocyanine photoacid **PAH**, the structure of which is depicted in Scheme 2. Upon exposure to violet light, **PAH** undergoes a ring-closing reaction to give the spiropyran **PA**[−], along with liberation of a proton (Scheme 2).¹¹ The spiropyran form is metastable and accumulates under irradiation. As a result, a pH drop of more than 2 units can be achieved. When light irradiation is stopped, the merocyanine form **PAH** is regenerated.

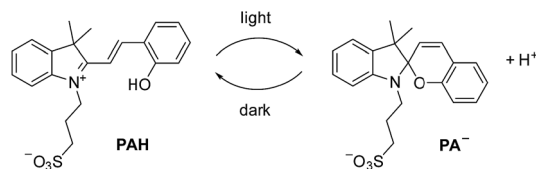
The metastable-state photoacid **PAH** was introduced by Liao and co-workers in 2011.¹² Since then, **PAH** and structurally related compounds have been used in the context of supramolecular chemistry. Applications include the light-induced complexation of a pyridinium salt by a macrocyclic receptor,¹³ the switching of a rotaxane-based molecular shuttle,¹⁴ the light-controlled reversible self-assembly of nanorods,¹⁵ the reversible chromism of guests in coordination cages,¹⁶ the modulation of a hydrazine switch,¹⁷ the reversible photo-induced gel–sol transition of a dipeptide gel,¹⁸ and the controlled aggregation of



Scheme 1 Reversible disassembly of metallasupramolecular structures by means of a metastable-state photoacid (**PAH**). The M₆L₁₂ coordination cage is used as a representative example of a metal–ligand assembly.

Institut des Sciences et Ingénierie Chimiques, Ecole Polytechnique Fédérale de Lausanne (EPFL), 1015 Lausanne, Switzerland. E-mail: kay.severin@epfl.ch

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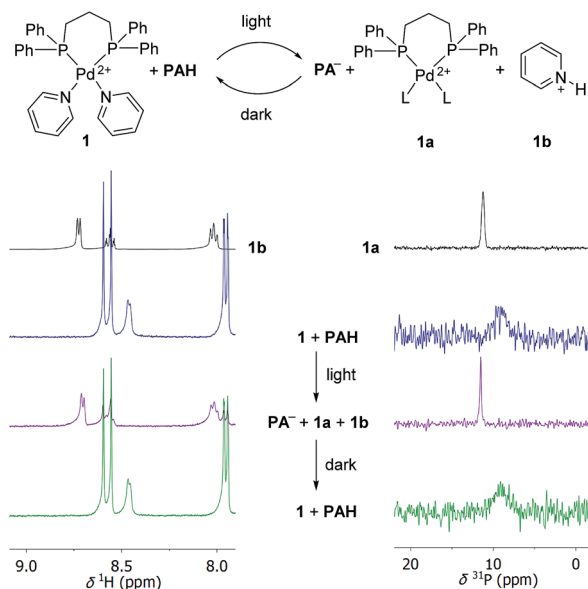


Scheme 2 Irradiation with light converts the merocyanine PAH into the spiropyran PA⁻, with concomitant liberation of a proton.

nanoparticles,¹⁹ among others.^{11,20} We wanted to explore if PAH is also suited for controlling the structures of metal–ligand assemblies by light.

First, we performed test reactions with the Pd^{II} complex [(dppp)Pd(py)₂](OTf)₂ (**1**). Complex **1** features a strongly bond 1,3-bis(diphenylphosphino)propane (dppp) ligand and two labile pyridine (py) ligands. We chose **1**, because Pd^{II} complexes with chelating diphosphine ligands are popular building blocks in metallasupramolecular chemistry.^{2,21} As connecting units, polypyridyl ligands are often employed, and the pyridyl ligands in **1** are surrogates for more intricate ligands. A solution of complex **1** (125 μM) and PAH (8 equiv.) was prepared in the dark using a mixture of CD₃CN and D₂O (8 : 2) as the solvent. The solution was then irradiated for 20 minutes with violet light using a commercial LED panel (λ = 425 nm). After irradiation, the sample was immediately analysed by ¹H and ³¹P NMR spectroscopy, and further NMR spectra were recorded after 8 hours in the dark.

The ¹H NMR spectrum of the mixture before irradiation shows two doublets in the region between 7.5 and 9 ppm, which can be assigned to the photoacid PAH (Scheme 3, left side). In addition, one can observe a doublet at ~8.5 ppm, resulting from



Scheme 3 Irradiation of a mixture of complex **1** (125 μM) and the photoacid PAH (8 equiv.) in CD₃CN/D₂O (8 : 2) by violet light results in the displacement of the pyridyl ligands as evidenced by the ¹H NMR (left side) and ³¹P NMR spectra (right side). The triflate anions of **1** are not shown for clarity. L indicates solvent or OTf⁻.

the pyridine NCH protons. Irradiation leads to the nearly complete disappearance of the two PAH signals. The NCH signal of the bound pyridine ligands is no longer observable. Instead, three new signals appear. These peaks can be attributed to the formation of protonated pyridine (**1b**), as evidenced by comparison with a reference sample of pyridinium chloride (Scheme 3, spectrum on the top left side). After 8 hours in the dark, the NMR spectrum of the sample is indistinguishable from that of the sample before irradiation. The ³¹P NMR spectrum of the starting mixture shows a broad peak at ~9 ppm, which is converted to a sharper signal at ~12 ppm after irradiation (Scheme 3, right side). A similar signal is observed when complex [(dppp)Pd](OTf)₂ is dissolved in CD₃CN/D₂O. We attribute this signal to a complex [(dppp)PdL₂] (**1a**), with L being either solvent and/or triflate. Again, the spectrum converts back to that of the starting mixture when the sample is kept in the dark. Taken together, the data are strong evidence for a light-induced displacement of the pyridine ligands, which is fully reversed in the dark.

Similar test reactions were performed with the model complexes [(py)₄Pd](BF₄)₂ (**2**), [(bipy)₃Fe](BF₄)₂ (**3**), and [(bipy)₃Zn](BF₄)₂ (**4**) (bipy = 2,2'-bipyridyl). These kind of coordination motifs are also found in numerous metallasupramolecular assemblies.²² As it was observed for **1**, it was possible to dissociate the N-donor ligands by irradiation with light, and re-complexation was observed in the dark (for details see the ESI†). For the Pd complex **2** and the Zn complex **4**, the photo-induced conversion was very high (>80%), whereas for the Fe complex **3**, only partial disassembly was observed (~46% of complex **3** remained intact). We have also attempted to break complex **3** with a Brønsted acid, namely trifluoroacetic acid (TFA, 24 equiv. with respect to Fe). As with PAH, only partial disassembly was observed, indicating that the Fe complex is less susceptible to an acid-induced ligand displacement.

Encouraged by the results with the model complexes **1–4**, we next examined reactions with larger and structurally more complex metal–ligand assemblies. Six different

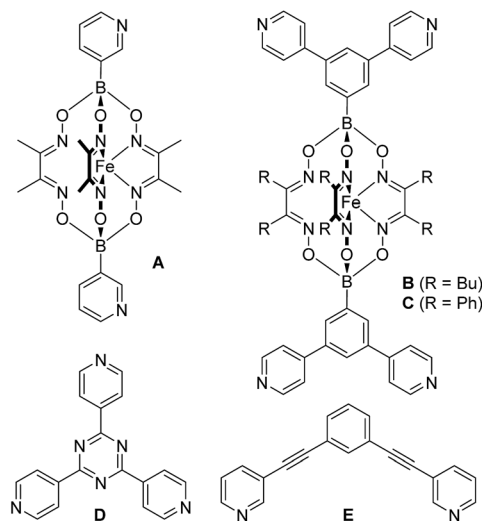


Fig. 1 Structures of the pyridyl ligands A–E.

metallasupramolecular structures were prepared using the pyridyl ligands depicted in Fig. 1. More specifically, we have prepared an octahedral M_6L_{12} cage with $[Pd(py^*)_4]^{2+}$ (py^* = substituted pyridine) complexes at the vertices, bridging ligands **A**, and a $[B(p-C_6H_4F)_4]^-$ guest molecule (**5**),²³ two M_6L_4 cages with gyrobifastigium-like geometry and $[(dppp)Pd(py^*)]^{2+}$ or $[(dcpm)Pt(py^*)]^{2+}$ complexes connecting the tetratopic ligand **B** (**6** and **7**, $dcpm$ = bis(dicyclohexylphosphino)methane),²⁴ a coordination barrel containing $[(dcpm)Pt(py^*)]^{2+}$ complexes and ligand **C** (**8**),^{24a} a trigonal prismatic Ru complex with tris-pyridyl ligand **D** (**9**),²⁵ and an M_2L_4 cage based on ligand **E** with $[Pd(py^*)_4]^{2+}$ vertices (**10**)²⁶ (Fig. 2).

First test reactions with complex **5** and TFA (24 equiv.) showed that the complex can be fully disassembled by acid. We then set up an experiment similar to what has been described for the model complexes. Assembly **5** was dissolved together with an excess of the photoacid **PAH** in a mixture of CD_3CN and D_2O (8 : 2) in the dark. After irradiation with violet light, we immediately recorded 1H and ^{19}F NMR spectra, followed by

additional spectra after a waiting period of 8 hours in the dark. Preliminary studies revealed that the addition of small amounts of NaCl facilitated the disassembly process, and all subsequent experiments with **5** were performed with six equivalents of NaCl. It is worth noting that chloride ions are known for their ability to decompose Pd^{II} cages.²⁷ For complex **5**, six equivalents of NaCl were not sufficient to induce detectable decomposition, even though we were able to disrupt the assembly at higher NaCl concentrations. The beneficial effect of small amounts of NaCl for the photoswitching process might be due to stabilization of the Pd^{II} ions after ligand displacement.

The 1H NMR spectrum after irradiation indicates the acid-induced de-complexation of the pyridyl ligands (Scheme 4, left side). Additional evidence for the disassembly of the octahedron was obtained by analysing the ^{19}F NMR spectra (Scheme 4, right side). The initial spectrum shows two signals, one for the encapsulated $[B(p-C_6H_4F)_4]^-$ guest molecule, and one for non-bound $[B(p-C_6H_4F)_4]^-$. After irradiation, only 'free' $[B(p-C_6H_4F)_4]^-$ (**5c**) is observed. In the dark, the M_6L_{12} cage is re-formed, as corroborated by the 1H NMR spectrum, and the ^{19}F NMR signal of the encapsulated borate anion. Overall, it is evident that light irradiation leads to a destruction of the metallasupramolecular structure and liberation of the guest. The process is reversible: in the dark, the cage with the encapsulated guest is formed again.

Experiments with the assemblies **6–10** gave similar results: a reversible, light-induced disassembly of the cage structures

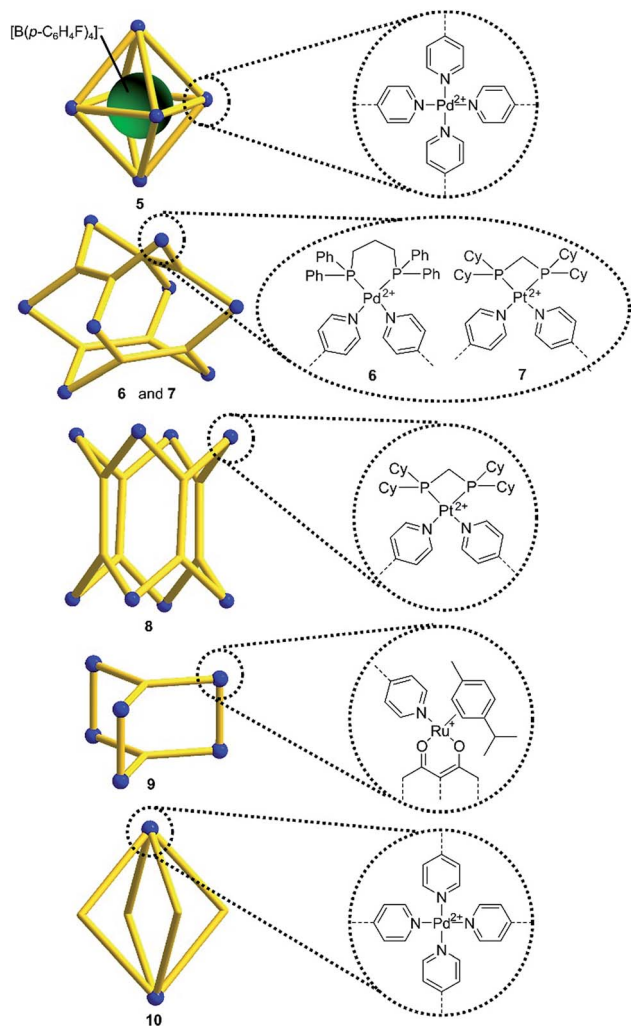
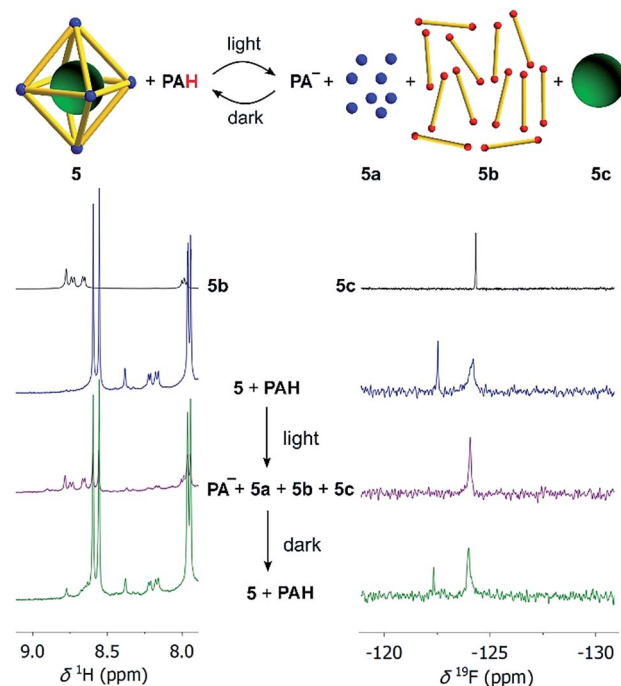


Fig. 2 Schematic representation of the metallasupramolecular structures used in this study. Details about the bridging ligands (yellow) are given in ESI.†



Scheme 4 Irradiation of a mixture of cage **5** (10 μM), NaCl (60 μM), and the photoacid PAH (8 equiv. with respect to each bipyridyl ligand) in CD_3CN/D_2O (8 : 2) by violet light results in disassembly of the cage and liberation of the $[B(p-C_6H_4F)_4]^-$ guest (green ball) as evidenced by the 1H NMR (left) and ^{19}F NMR spectra (right). Details about the bridging ligand (yellow) are given in the ESI.†

was observed by NMR spectroscopy in all cases (for details see the ESI†). For **6**, **7**, and **8**, the switching efficiency was found to be very high, with at least 80% of the cages being disassembled by the photoacid. For **9** and **10**, the process was less efficient, with only 16% (**9**) and 45% (**10**) of the cage structures being disrupted.

The difference in switching efficiency between cage **5** (>80%) and cage **10** (45%) is intriguing, given that both assemblies are based on $[\text{Pd}(\text{py}^*)_4]^{2+}$ complexes. We were able to identify two factors which contribute to the reduced switching efficiency observed for **10**: first, the photoacid binds weakly to cage **10** (Fig. S31†),²⁸ and this interaction is expected to stabilize the assembly. Such an interaction was not observed for cage **5**. Second, the pyridyl ligand **A** used for cage **5** is more basic than the pyridyl ligand **E** used for cage **10** (Fig. S32†). Consequently, acid-induced disassembly is favoured for cage **5**.

The disassembly process was fully reversible in all cases. Kinetic investigations at 298 K showed that all assemblies are re-formed with half-lives between 20 and 70 minutes. For comparison: the photoacid alone switches back with $t_{1/2}$ of 2.4 minutes under these conditions ($\text{CH}_3\text{CN}/\text{H}_2\text{O}$, 8 : 2), indicating that the re-formation of the coordination complex is the rate-limiting step. The re-assembly of the complexes **5** and **8** was also followed at 50 °C. As expected, the reactions were faster, and complete re-assembly was observed after 30 minutes (**5**) or 10 minutes (**8**), respectively.

The robustness of the photoswitching process was demonstrated by repeated disassembly–assembly cycles using the coordination barrel **8** (Fig. 3). A mixture of the barrel and **PAH** was irradiated for 20 minutes, and then ^1H NMR and ^{31}P NMR spectra were recorded. After 2.5 hours in the dark, additional spectra were measured. This procedure was repeated four times. The ^1H and ^{31}P NMR spectra show barrel destruction and re-assembly over all cycles. The relative amount of the protonated ligand **8b**, as determined by ^1H NMR spectroscopy, was used as an indication of the switching efficiency. The process displayed good stability over the five cycles (Fig. 3). Ultimately, the sensitivity of **PAH** towards hydrolytic degradation represents a limit,^{12,13} but within the timeframe of our experiment, hydrolysis was not observed.

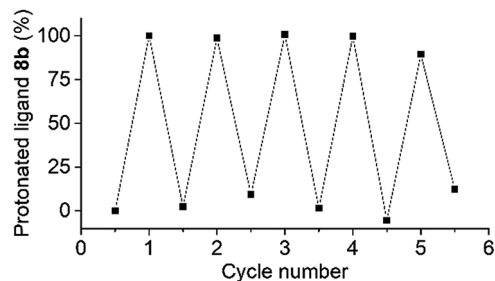


Fig. 3 Repeated photoswitching between the coordination barrel **8** and its disassembled state. The relative amount of the protonated ligand **8b** was used as an indication of the switching efficiency. For the first cycle, the amount of **8b** was normalized to 0 and 100%, respectively.

Conclusions

We have demonstrated that the addition of a metastable-state photoacid to solutions of metallasupramolecular assemblies renders the systems photo-responsive. Irradiation with violet light leads to proton transfer from the photoacid to the ligand, resulting in partial or full disassembly of the metal–ligand structure. The process is in all cases reversed in the dark. In principle, similar changes could be induced by addition of acids and bases. However, the utilization of light is not effecting the overall composition of the system (*e.g.* no accumulation of salts), and light can be used with good spatial and temporal control. Compared to the utilization of photochromic ligands, our approach offers the advantage of being more general and easier to implement. For the present study, we have focused on metallasupramolecular structures based on polypyridyl ligands, but it is expected that the methodology can be extended to assemblies with other ligands, given that the basicity of the system is within the pH range of the photoacid. We have already shown that the approach is suited for the light-induced release of guest molecules, but other applications can be envisioned as well. For example, it could be possible to control catalytic reactions with light, since some coordination cages are able to act as catalytic nanoreactors.²⁹ Furthermore, it should be possible to use a related approach for photoswitching of metallasupramolecular polymers and metallogels.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

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References

- (a) A. Díaz-Moscoco and P. Ballester, *Chem. Commun.*, 2017, 53, 4635–4652; (b) M. Kathan and S. Hecht, *Chem. Soc. Rev.*, 2017, 46, 5536–5550; (c) X. Yao, T. Li, J. Wang, X. Ma and H. Tian, *Adv. Opt. Mater.*, 2016, 4, 1322–1349; (d) D.-H. Qu, Q.-C. Wang, Q.-W. Zhang, X. Ma and H. Tian, *Chem. Rev.*, 2015, 115, 7543–7588; (e) S. Yagai and A. Kitamura, *Chem. Soc. Rev.*, 2008, 37, 1520–1529.
- (a) W. Wang, Y.-X. Wang and H.-B. Yang, *Chem. Soc. Rev.*, 2016, 45, 2656–2693; (b) A. J. McConnell, C. S. Wood, P. P. Neelakandan and J. R. Nitschke, *Chem. Rev.*, 2015, 115, 7729–7793.
- M. Han, Y. Luo, B. Damaschke, L. Gómez, X. Ribas, A. Jose, P. Peretzki, M. Seibt and G. H. Clever, *Angew. Chem., Int. Ed.*, 2016, 55, 445–449.
- M. Han, R. Michel, B. He, Y.-S. Chen, D. Stalke, M. John and G. H. Clever, *Angew. Chem., Int. Ed.*, 2013, 52, 1319–1323.



- 5 S.-C. Wei, M. Pan, Y.-Z. Fan, H. Liu, J. Zhang and C.-Y. Su, *Chem.-Eur. J.*, 2015, **21**, 7418–7427.
- 6 T. Murase, S. Sato and M. Fujita, *Angew. Chem., Int. Ed.*, 2007, **46**, 5133–5136.
- 7 J. Park, L.-B. Sun, Y.-P. Chen, Z. Perry and H.-C. Zhou, *Angew. Chem., Int. Ed.*, 2014, **53**, 5842–5846.
- 8 The solubility of a coordination cage can also be modified with a photoswitchable guest. See: G. H. Clever, S. Tashiro and M. Shionoya, *J. Am. Chem. Soc.*, 2010, **132**, 9973–9975.
- 9 D. Zhang, Y. Nie, M. L. Saha, Z. He, L. Jiang, Z. Zhou and P. J. Stang, *Inorg. Chem.*, 2015, **54**, 11807–11812.
- 10 For examples of acid/base-switchable supramolecular assemblies see: (a) J. Altmann and A. Pöthing, *Angew. Chem., Int. Ed.*, 2017, **56**, 15733–15736; (b) S. Yi, V. Brega, B. Captain and A. E. Kaifer, *Chem. Commun.*, 2012, **48**, 10295–10297; (c) S. Hiraoka, Y. Sakata and M. Shionoya, *J. Am. Chem. Soc.*, 2008, **130**, 10058–10059; (d) P. Mal, D. Schultz, K. Beyeh, K. Rissanen and J. R. Nitschke, *Angew. Chem., Int. Ed.*, 2008, **47**, 8297–8301; (e) T. Gottschalk, B. Jaun and F. Diederich, *Angew. Chem., Int. Ed.*, 2006, **46**, 260–264; (f) Z. Grote, R. Scopelliti and K. Severin, *J. Am. Chem. Soc.*, 2004, **126**, 16959–16972; (g) F. Ibukuro, T. Kusukawa and M. Fujita, *J. Am. Chem. Soc.*, 1998, **120**, 8561–8856; (h) N. Branda, R. M. Grotzfeld, C. Valdes and J. Rebek, *J. Am. Chem. Soc.*, 1995, **117**, 85–88.
- 11 Y. Liao, *Acc. Chem. Res.*, 2017, **50**, 1956–1964.
- 12 Z. Shi, P. Peng, D. Strohecker and Y. Liao, *J. Am. Chem. Soc.*, 2011, **133**, 14699–14703.
- 13 Q. Shi and C.-F. Chen, *Org. Lett.*, 2017, **19**, 3175–3178.
- 14 L.-P. Yang, F. Jia, J.-S. Cui, S.-B. Lu and W. Jiang, *Org. Lett.*, 2017, **19**, 2945–2948.
- 15 J. Guo, H. Zhang, Y. Zhoua and Y. Liu, *Chem. Commun.*, 2017, **53**, 6089–6092.
- 16 D. Samanta, D. Galaktionova, J. Gemen, L. J. W. Shimon, Y. Diskin-Posner, L. Avram, P. Král and R. Klajn, *Nat. Commun.*, 2018, **9**, 641.
- 17 L. A. Tatum, J. T. Foy and I. Aprahamian, *J. Am. Chem. Soc.*, 2014, **136**, 17438–17441.
- 18 X. Li, J. Fei, Y. Xu, D. Li, T. Yuan, G. Li, C. Wang and J. Li, *Angew. Chem., Int. Ed.*, 2018, **130**, 1921–1925.
- 19 P. K. Kundo, D. Samanta, R. Leizrowice, B. Margulis, H. Zhao, M. Börner, T. Udayabhaskararao, D. Manna and R. Klajn, *Nat. Chem.*, 2015, **7**, 646–652.
- 20 For an early study on this topic see: S. Silvi, A. Arduini, A. Pochini, A. Secci, M. Tamasulo, F. M. Raymo, M. Baroncini and A. Credi, *J. Am. Chem. Soc.*, 2007, **129**, 13378–13379.
- 21 (a) T. R. Cook and P. J. Stang, *Chem. Rev.*, 2015, **115**, 7001–7045; (b) R. Chakrabarty, P. S. Mukherjee and P. J. Stang, *Chem. Rev.*, 2011, **111**, 6810–6918.
- 22 (a) Z. Wu, K. Zhou, A. V. Ivanov, M. Yusobov and F. Verpoort, *Coord. Chem. Rev.*, 2017, **353**, 180–200; (b) L.-J. Chen, H.-B. Yang and M. Shionoya, *Chem. Soc. Rev.*, 2017, **46**, 2555–2576; (c) W. M. Bloch and G. H. Clever, *Chem. Commun.*, 2017, **53**, 8506–8516; (d) M. Han, D. M. Engelhard and G. H. Clever, *Chem. Soc. Rev.*, 2014, **43**, 1848–1880; (e) S. Mukherjee and P. S. Mukherjee, *Chem. Commun.*, 2014, **50**, 2239–2248.
- 23 M. D. Wise, J. J. Holstein, P. Pattison, C. Besnard, E. Solari, R. Scopelliti, G. Bricogne and K. Severin, *Chem. Sci.*, 2015, **6**, 1004–1010.
- 24 (a) G. Cecot, M. Marmier, S. Geremia, R. De Zorzi, A. V. Vologzhanina, P. Pattison, E. Solari, F. F. Tirani, R. Scopelliti and K. Severin, *J. Am. Chem. Soc.*, 2017, **139**, 8371–8381; (b) G. Cecot, B. Alameddine, S. Prior, R. De Zorzi, S. Geremia, R. Scopelliti, F. T. Fadaei, E. Solari and K. Severin, *Chem. Commun.*, 2016, **52**, 11243–11246.
- 25 N. P. E. Barry and B. Therrien, *Eur. J. Inorg. Chem.*, 2009, 4695–4700.
- 26 D. P. August, G. S. Nichol and P. J. Lusby, *Angew. Chem., Int. Ed.*, 2016, **55**, 15022–15026.
- 27 (a) D. Preston, A. Fox-Charles, W. K. C. Lo and J. D. Crowley, *Chem. Commun.*, 2015, **51**, 9042–9045; (b) R. Zhu, J. Lübben, B. Dittrich and G. H. Clever, *Angew. Chem., Int. Ed.*, 2015, **54**, 2796–2800; (c) J. E. M. Lewis, E. L. Gavey, S. A. Cameron and J. D. Crowley, *Chem. Sci.*, 2012, **3**, 778–784.
- 28 For examples of Pd^{II} cages which bind guest molecules with sulfonate groups see: (a) S. Löffler, J. Lübben, A. Wuttke, R. A. Mata, M. John, B. Dittrich and G. H. Clever, *Chem. Sci.*, 2016, **7**, 4676–4684; (b) G. H. Clever and M. Shionoya, *Chem.-Eur. J.*, 2010, **16**, 11792–11796; (c) S. Hiraoka, M. Kiyokawa, S. Hashida and M. Shionoya, *Angew. Chem., Int. Ed.*, 2010, **49**, 138–143.
- 29 (a) M. Otte, *ACS Catal.*, 2016, **6**, 6491–6510; (b) L. Catti, Q. Zhang and K. Tiefenbacher, *Chem.-Eur. J.*, 2016, **22**, 9060–9066.

