

Facile preparation, high microwave absorption and microwave absorbing mechanism of RGO-Fe₃O₄ composites†

Cite this: *RSC Adv.*, 2013, **3**, 23638

Meng Zong, Ying Huang,* Yang Zhao, Xu Sun, Chunhao Qu, Didi Luo and Jiangbo Zheng

Reduced graphene oxide (RGO)-Fe₃O₄ composites with obviously enhanced microwave absorption properties were successfully fabricated by a rational one-pot simplified co-precipitation route, which avoided the usage of an inert gas and any additional chemical agents (such as surfactants and stabilizers). Given these advantages, the strategy described in this study can be developed as a simple and large-scale route to yield RGO-Fe₃O₄ composites. The morphology, structure, thermal stability, magnetic and microwave electromagnetic properties of the as-prepared composites were characterized by XRD, XPS, TEM, FT-IR, Raman, TG and VSM. These composites exhibit excellent microwave absorption properties, which are attributed to effective complementarities between the dielectric loss and the magnetic loss. The microwave absorption mechanism of the RGO-Fe₃O₄ composites was studied in detail. For the RGO-Fe₃O₄-3 composite, the maximum R_L reaches -44.6 dB at 6.6 GHz with a thickness of 3.9 mm, and the bandwidth of R_L less than -10 dB can reach up to 4.3 GHz (from 12.2 to 16.5 GHz) with a thickness of 2.0 mm. Moreover, the microwave absorption properties can be tuned easily by varying the (RGO)/(Fe₃O₄) ratio and layer thickness of the samples. It is believed that such composites will find wide applications in the microwave absorbing area.

Received 2nd July 2013

Accepted 9th September 2013

DOI: 10.1039/c3ra43359e

www.rsc.org/advances

1. Introduction

With the development of electronic technology, electromagnetic (EM) interference has become a pollution problem, not only influencing the operation of electronic devices, but also being harmful to the health of human beings.^{1,2} Microwave absorbing materials have attracted much attention owing to these expanding EM interference problems.³⁻⁵ They can absorb microwaves effectively and convert EM energy into thermal energy or dissipate microwaves by interference. Microwave absorbing materials are now required to have strong absorption characteristics, wide absorption frequencies, be lightweight and have antioxidation properties.⁴ There are a number of microwave absorbing materials such as ferrites, carbonyl iron, conducting polymers and carbon-based materials *etc.*⁶⁻⁹ However, the traditional microwave absorbing materials cannot meet all of the requirements such as to be strong, wide, lightweight and thin at the same time. Hence, extensive studies have been made to develop novel microwave absorbing

materials with high absorption and wide absorption frequencies.

Graphene, a novel carbon nanomaterial consisting of one-atom-thick, hexagonally arranged carbon atoms, has attracted extensive attention owing to its superior electronic, thermal and mechanical properties as well as its chemical stability.^{10,11} Due to its special surface properties and layered structure, graphene has become a potential nanoscale building block for new hybrid materials.¹² Recent research shows that inorganic nanoparticles, such as Au, Ag, Fe₃O₄ and Co₃O₄, *etc.*, could be attached to graphene or graphene oxide (GO) to form hybrid materials, which have potential applications in surface enhanced Raman scattering (SERS), lithium ion batteries, hydrogen storage, microwave absorption and as a magnetic resonance imaging (MRI) contrast agent, *etc.*¹³⁻¹⁶

As a class of promising microwave absorbing materials, carbon-based materials exhibited several exceptional properties including being lightweight, having wide absorption frequency, high thermal stability, and high chemical stability.¹ In previous research, Zhan and co-workers^{17,18} did some work on the synthesis of the heterojunction of Fe₃O₄ with carbon nanotubes (CNTs) or graphene, revealing that hybrid materials can reinforce the original properties and extend the applications of simplex materials. In comparison with CNTs, graphene possesses similar physical properties but has larger surface areas, a lower price and more stable properties. Furthermore,

Department of Applied Chemistry, The Key Laboratory of Space Applied Physics and Chemistry, Ministry of Education, School of Science, Northwestern Polytechnical University, Xi'an 710072, PR China. E-mail: mwpu_zongmeng@163.com; yingh@mwpu.edu.cn; Tel: +86 29 88431636

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c3ra43359e

more effective absorption could be obtained by platelet-shaped materials than by the rod-shaped and sphere-shaped ones in microwave absorption applications.¹⁷ Therefore, graphene is highly expected to be the alternative to CNTs in microwave absorbing materials.

As is known, Fe₃O₄ has an inverse spinel structure, in which half of the Fe³⁺ cations occupy tetrahedral (A) sites, while all of the Fe²⁺ cations and the other half of the Fe³⁺ cations occupy octahedral (B) sites, resulting in the structural formula of Fe³⁺_AFe³⁺Fe²⁺_BO₄.¹⁹ It is an important magnetic material and has physical and chemical properties such as superparamagnetism, half-metallic character and strong spin polarization at room temperature in addition to its low toxicity and high biocompatibility. Moreover, its magnetic properties can be tuned by size, shape, and dimension.⁴ Thus, Fe₃O₄ has wide applications in many areas such as gas sensors, magnetic storage media, optoelectronic and spintronic devices, biomedicine, as well as materials for microwave absorbing and shielding research.^{4,20–22} However, its heavy weight restrain it from being widely used in the field of electromagnetic wave absorption.²³ Furthermore, its microwave absorption properties are usually subject to the degradation caused by the eddy current in the high-frequency region.⁴

Herein, the reduced graphene oxide (RGO) coated with Fe₃O₄ nanoparticles (NPs) was synthesized through a facile one-pot simplified co-precipitation method, which avoided the usage of an inert gas or any additional surfactants or stabilizers. The strategy described in this study can be developed as a simple and large-scale route to yield RGO–Fe₃O₄ composites. The crystalline structure, morphology, thermal stability, magnetic properties and microwave electromagnetic properties of the as-prepared composites were investigated. The mechanism of the enhanced reflection loss for the RGO–Fe₃O₄ composites will also be explained comprehensively.

2. Experimental method

2.1 Preparation of RGO–Fe₃O₄ composites

GO was prepared by Hummers' method as reported in ref. 24. The composites of RGO–Fe₃O₄ were prepared by a one-pot simplified co-precipitation route. In a typical procedure, 100 mg GO was dispersed in 150 mL of deionized water and ultrasonicated for 2 h. 88 mg FeCl₃·6H₂O and 65 mg FeCl₂·4H₂O (Fe³⁺ and Fe²⁺ with a mole ratio of 1 : 1) were dispersed in 50 mL deionized water and were then added to the suspension of GO. 1 M NaOH aqueous solution was added to the suspension until the pH = 11, and the mixture was stirred for 2 h at 80 °C without protection of an inert gas. Subsequently, a freshly prepared NaBH₄ aqueous solution (NaBH₄ and GO with a mass ratio of 6 : 1) was added dropwise with stirring and the mixture was stirred again for 1 h at 80 °C to ensure complete reduction of GO. The black products were washed several times with deionized water and ethanol, then dried at 60 °C under vacuum and denoted as RGO–Fe₃O₄-1. With different weights of GO (50 mg, 25 mg, 0 mg) being added to the reaction system, the final samples were labeled as RGO–Fe₃O₄-2, RGO–Fe₃O₄-3, and Fe₃O₄, respectively. For comparison purposes, RGO was also

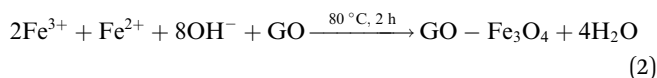
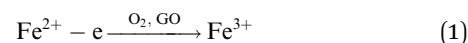
prepared in similar procedures in the absence of Fe³⁺, Fe²⁺ and NaOH.

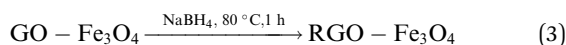
2.2 Characterization

The crystal structure was determined by X-ray diffraction (XRD, Rigaku, model D/max-2500 system at 40 kV and 100 mA of Cu K α). XPS analysis was characterized by an X-ray photoelectron spectrometer (K-Alpha; Thermo Fisher Scientific (SID-Elemental), New York, USA). The morphology and the size of the synthesized samples were characterized by transmission electron microscopy (TEM, Tecnai F30 G2, FEI, USA). The Fourier transform infrared spectroscopy (FT-IR) spectra of the composites were obtained by using a Model NicoLETiS10 Fourier transform spectrometer (Thermo SCIENTIFIC Co., USA) with a 2 cm⁻¹ resolution in the range of 4000–400 cm⁻¹. The Raman spectra of the composites were obtained by using an in Via Laser Raman spectrometer (Renishaw Co., England) with 514 nm radiation. The thermal stabilities of the composites were analyzed by using thermogravimetric analysis (TGA, Model Q50, TA, USA) from room temperature to 800 °C in an air atmosphere, with a heating rate of 20 °C min⁻¹. The magnetic properties were measured by a vibrating sample magnetometer (VSM, Riken Denshi, BHV-525) at room temperature. Electromagnetic (EM) parameters were measured by a vector network analyzer (NA, HP8720ES) in the range of 2–18 GHz, in which the synthesized powders were pressed to be toroidal samples (outer diameter: 7 mm, inner diameter: 3.04 mm, and thickness: 3 mm) according to the mass ratio 1 : 1 of paraffin to RGO–Fe₃O₄ composite.

3. Results and discussions

The formation of the RGO, Fe₃O₄ NPs and RGO–Fe₃O₄ composites is described in Fig. S1 (please refer to ESI†). The graphite was treated with H₂SO₄ and KMnO₄ and GO was obtained which contained a variety of functional groups including –COOH, –OH, epoxy and ketone.^{17,25} Co-precipitation is a common method to synthesize RGO–Fe₃O₄ composites. This method always uses an inert gas (such as N₂ or Ar) to prevent the oxidation of Fe²⁺ by O₂ when Fe²⁺ : Fe³⁺ is 1 : 2.²⁵ In order to avoid the usage of an inert gas, we increased the ratio of the Fe²⁺. In our work, Fe³⁺ and Fe²⁺ were dispersed in a suspension of GO with a molar ratio of 1 : 1 and adsorbed onto the surface of the GO by an electrostatic attraction. Part of the Fe²⁺ cations were oxidized by O₂ and GO.²⁶ Then the iron ions were transformed into Fe₃O₄ NPs on the surface of the GO by addition of NaOH at elevated temperatures. Therefore, there may be a C–O–Fe bond between the RGO and Fe₃O₄ (Fig. S2†). Finally the GO–Fe₃O₄ was reduced to RGO–Fe₃O₄ composites by NaBH₄. The overall reaction can be described by the following equations:





The morphologies and structures of the RGO and RGO-Fe₃O₄ composites were investigated by transmission electron microscopy (TEM), which is shown in Fig. 1 and S1.† Fig. 1a and the inset in Fig. 1a show the low and high magnification TEM image of RGO, respectively, and we can see that the RGO nanosheets are less than 3 nm thick. The RGO sheets from the reduction by NaBH₄ of GO are transparent thin films (Fig. S1b†), and comparing with GO (Fig. S1a†), the RGO sheets have a crumpled and rippled structure which is due to deformation during the exfoliation and restacking process. Fig. 1b–d and S1c–e† show the low magnification TEM images of the obtained RGO-Fe₃O₄ composites with different (RGO)/(Fe₃O₄) ratios. It is clear that the RGO nanosheets are well loaded by Fe₃O₄ NPs, and the diameters of the Fe₃O₄ NPs are 15–25 nm. The mass ratio of (RGO)/(Fe₃O₄) has an important effect on the loading density of the Fe₃O₄ NPs on the RGO sheets. As shown in Fig. 1b–d, the densities of the Fe₃O₄ NPs increase gradually with the decrease of the GO amount used. This result is due to the fact that less GO will take up comparatively more Fe³⁺ ions to

nucleate, thus leading to the increase in density of Fe₃O₄ NPs. Therefore, the amount of GO can effectively tune the density of the Fe₃O₄ NPs.²³ Moreover, almost no Fe₃O₄ NPs are found outside of the RGO nanosheets, indicating that the simplified co-precipitation synthesis of RGO-Fe₃O₄ composites has a high efficiency. The HRTEM image shown in Fig. 1e also reveals the crystalline structure of the Fe₃O₄ NPs, and the lattice fringes with interplanar distances of 0.258 nm can be assigned to the (311) planes of the cubic spinel crystal Fe₃O₄. The selected-area electron diffraction pattern (SAED) (Fig. 1f) clearly shows a ring pattern arising from the cubic spinel crystal Fe₃O₄, further confirming the crystalline nature of the Fe₃O₄ NPs.²⁷

Powder X-ray diffraction (XRD) measurements were utilized to investigate the phase composition and the crystalline structure of the synthesized samples. As shown in Fig. 2a, the XRD pattern of GO shows a sharp peak at $2\theta = 11.1^\circ$, corresponding to the (001) reflection of GO. The diffraction peak of the as-prepared RGO (Fig. 2b) at 25.6° could be attributed to the graphitic structure (002) of the short-range order in stacked graphene sheets.²⁸ After surface decoration (Fig. 2c–e), six new peaks are observed, which are very similar to that of the pure Fe₃O₄ NPs (Fig. 2f), and could be indexed as the characteristic (220), (311), (400), (422), (511) and (440) reflections of the pure cubic spinel crystal structure of Fe₃O₄ (JCPDS no. 19-0629).²⁹ It can be seen that except for the peaks assigned to Fe₃O₄, no other diffraction peaks resulting from GO or graphene can be found, which indicates that the GO is effectively reduced into graphene and the self-restacking of the as-reduced graphene sheets is effectively prevented.¹³ In fact, a similar result is also obtained and reported in ref. 16, 23 and 30. No peaks corresponding to impurities are detected, therefore a RGO-Fe₃O₄ heterostructure has been formed *via* a simplified co-precipitation method.

The XRD patterns of magnetite Fe₃O₄ (JCPDS no. 19-0629) and maghemite γ -Fe₂O₃ (JCPDS no. 39-1346) are highly similar, so X-ray photoelectron spectroscopy (XPS) measurements have to be consulted to unambiguously assign the crystal phase



Fig. 1 TEM images of RGO (a), RGO-Fe₃O₄-1 composite (b), RGO-Fe₃O₄-2 composite (c), RGO-Fe₃O₄-3 composite (d); HRTEM image (e) and SAED pattern (f) of RGO-Fe₃O₄-2 composite.

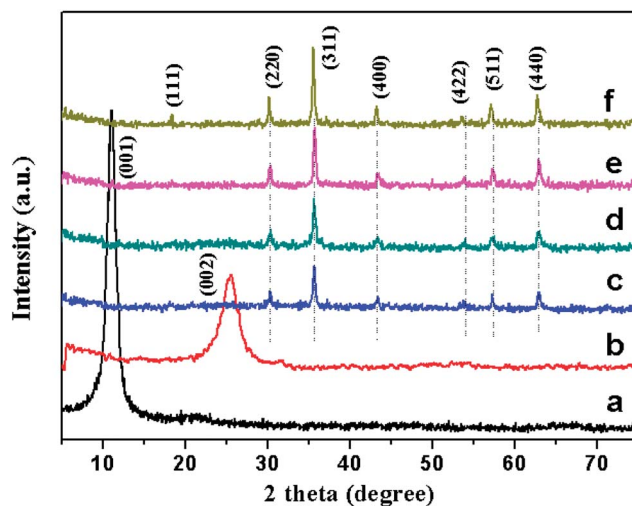


Fig. 2 XRD patterns of GO (a), RGO (b), RGO-Fe₃O₄-1 composite (c), RGO-Fe₃O₄-2 composite (d), RGO-Fe₃O₄-3 composite (e) and Fe₃O₄ NPs (f).

because XPS is very sensitive to Fe^{2+} and Fe^{3+} cations.^{23,31} Since we have adopted the same synthesis process for all composites, XPS was recorded only for Fe_3O_4 -RGO-1 as a representative and compared with that of GO.

Fig. 3a shows the wide scan spectrum of GO and the RGO- Fe_3O_4 -1 composite. The bands observed in the wide scan XPS spectrum of RGO- Fe_3O_4 -1 confirm the presence of C 1s, O 1s, and Fe 2p. The observed O 1s peak in GO at 533.1 eV is shifted to a lower binding energy (531.4 eV) due to the attachment of the Fe_3O_4 NPs in the Fe_3O_4 -RGO composites.¹⁶ As shown in Fig. 3b, the peaks located at 712.9 eV and 725.3 eV correspond to Fe 2p_{3/2} and Fe 2p_{1/2}, respectively,³² and there are no satellites for γ - Fe_2O_3 identified, excluding the presence of γ - Fe_2O_3 in the Fe_3O_4 samples,³¹ which confirms that the oxide in the sample is magnetite Fe_3O_4 . The C 1s spectrum of GO (Fig. 3c) and the Fe_3O_4 -RGO-1 composite (Fig. 3d) consist of four main components, arising from C=C/C-C (284.6 eV) in the aromatic rings, C-O (286.4 eV) of epoxy and alkoxy, C=O (287.8 eV) and O-C=O (289.3 eV) groups.³³ After reduction, it is clear that the C/O ratio in the RGO- Fe_3O_4 -1 composite increases remarkably compared with that of GO, and that most of the epoxide and hydroxyl functional groups are successfully removed, as shown in Fig. 3a, c and d. Such a higher O/C ratio of the RGO- Fe_3O_4 -1 composite implies good electronic conductivity, which may

enable the GO sheets to serve as conductive channels between the Fe_3O_4 NPs, and is favorable for microwave absorbing.

Fourier transform infrared spectroscopy (FT-IR) was used to indicate the degree of removal of the oxygen-containing

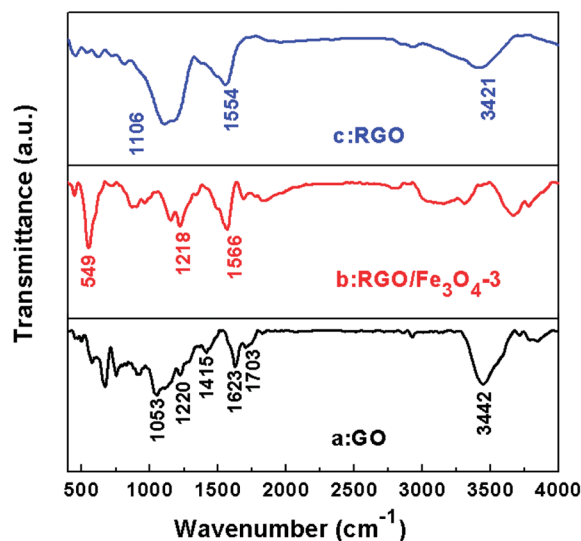


Fig. 4 FT-IR spectra of GO (a), the RGO- Fe_3O_4 -3 composite (b) and RGO (c).

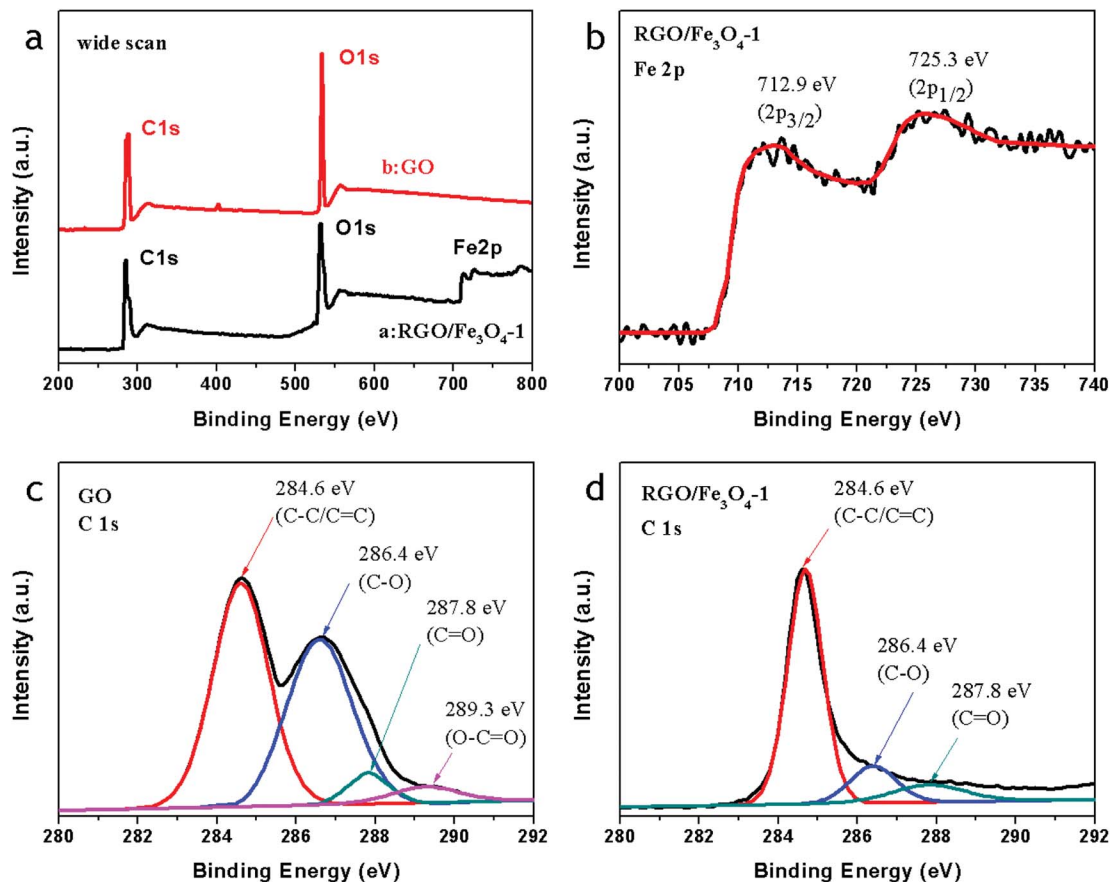


Fig. 3 X-ray photoelectron spectroscopy (XPS) spectra: wide scan (a), Fe 2p spectrum of the RGO- Fe_3O_4 -1 composite (b), C 1s spectrum of RGO (c), and C 1s spectrum of the RGO- Fe_3O_4 -1 composite (d).

functional groups. Fig. 4 shows the FT-IR transmittance spectra of GO (a), the RGO-Fe₃O₄-3 composite (b), and RGO (c). The spectra are shifted downward for easy viewing. The observed representative peaks in GO confirm the presence of the oxygen-containing functional groups in the carbon frameworks, which include bands at $\sim 1057\text{ cm}^{-1}$ (C-O stretching vibration), $\sim 1220\text{ cm}^{-1}$ (C-O stretching vibration of epoxide), $\sim 1415\text{ cm}^{-1}$ (O-H stretching vibration of carboxyl), $\sim 1703\text{ cm}^{-1}$ (C=O stretching of carbonyl and carboxyl groups at edges of the GO networks), while the band at $\sim 3420\text{ cm}^{-1}$ could be due to the O-H stretching mode of intercalated water.^{27,34-37} The band at $\sim 1623\text{ cm}^{-1}$ is assigned to skeletal vibrations of the unoxidized graphitic domains.^{16,27} All the absorption bands related to the oxygen-containing functional groups almost vanish in the RGO-Fe₃O₄-3 composite (Fig. 4b) and RGO (Fig. 4c), revealing that these oxygen-containing functional groups are removed in the presence of NaBH₄, and accordingly the GO is reduced into graphene. The band at *ca.* 1560 cm^{-1} in Fig. 4b and c could be ascribed to the formation of -COO⁻ after coating with the Fe₃O₄ NPs.¹⁶ Moreover, an additional peak at *ca.* 549 cm^{-1} in Fig. 4c can be ascribed to the lattice absorption of Fe₃O₄, further confirming the existence of Fe₃O₄ NPs.²⁷

Raman spectroscopy is a powerful nondestructive tool to distinguish ordered and disordered crystal structures of carbon.³⁴ Fig. 5 shows the Raman spectra of GO, RGO and the RGO-Fe₃O₄ composites. For the GO, graphene and graphene composites, there are always two peaks appearing at about 1340 cm^{-1} and 1590 cm^{-1} , which are known as the D-band and G-band.³⁸ The D band is a first-order zone boundary phonon mode associated with defects in the graphene or graphene edge, while the G band is a radial C-C stretching mode of sp² bonded carbon.^{34,39} A weak D band indicates a low density of defects and the intensity ratio of D band to G band (I_D/I_G) reflects the degree of the defects in the graphene or the edges.³⁹ Compared with GO, the ratio of intensities (I_D/I_G) for the RGO-Fe₃O₄ composites shows a remarkable increase, and the G band of the RGO-Fe₃O₄ composites becomes weaker and broader, suggesting a

higher degree of defects in the graphene or the edges due to the reduction of NaBH₄ and introduction of the Fe₃O₄ NPs.⁴⁰ The defects play an important role in microwave absorbing,⁴¹ and will further the interpretation of the microwave absorbing mechanism.

We have also examined the thermal stability of the prepared GO, RGO and RGO-Fe₃O₄ composites using thermogravimetry analysis (TGA). Thermal stability is an important material property of microwave absorption materials. Fig. 6 shows the TG analysis of GO, RGO and the RGO-Fe₃O₄ composites in an air atmosphere at a heating rate of $20\text{ }^\circ\text{C min}^{-1}$.

As shown in Fig. 6, the GO exhibits about 10 wt% loss below $100\text{ }^\circ\text{C}$ and more than 41 wt% loss at $240\text{ }^\circ\text{C}$, resulting from the removal of the labile oxygen-containing functional groups such as CO, CO₂, and H₂O.^{37,42} It shows a 5 wt% loss in an air atmosphere at $240\text{ }^\circ\text{C}$ for the RGO, which is much lower than that of the GO, indicating a decreased amount of oxygenated functional groups. A significant drop in mass of RGO around $600\text{ }^\circ\text{C}$ is due to the bulk pyrolysis of the carbon skeleton, which is similar to the drop in mass of GO around $550\text{ }^\circ\text{C}$. Thus, from the above analysis, the removal of the thermally labile oxygen functional groups by chemical reduction increased the thermal stability of graphene.³⁷

From the TG curves, we can see that the RGO-Fe₃O₄ composites display three weight loss processes.³⁰ The slight mass loss below $100\text{ }^\circ\text{C}$ is attributed to the evaporation of absorbed water. Then, a gradual weight loss occurs from $100\text{ }^\circ\text{C}$ to $350\text{ }^\circ\text{C}$, which can be assigned to the removal of the labile oxygen-containing functional groups and H₂O vapor from the sample caused by the destruction of oxygenated functional groups. Next a significant weight loss occurs between $350\text{ }^\circ\text{C}$ and $510\text{ }^\circ\text{C}$, indicating the oxidation and decomposition of graphene in air, leaving the final product as Fe₂O₃.³⁰ The skeleton of the graphene in the RGO-Fe₃O₄ composites decomposes at a lower temperature ($350\text{--}510\text{ }^\circ\text{C}$) compared to that of the RGO ($480\text{--}620\text{ }^\circ\text{C}$), which might be caused by the reaction of the skeleton carbon atoms and the iron oxide nanoparticles.⁴³



Fig. 5 Raman spectra for GO, RGO and RGO-Fe₃O₄ composites.

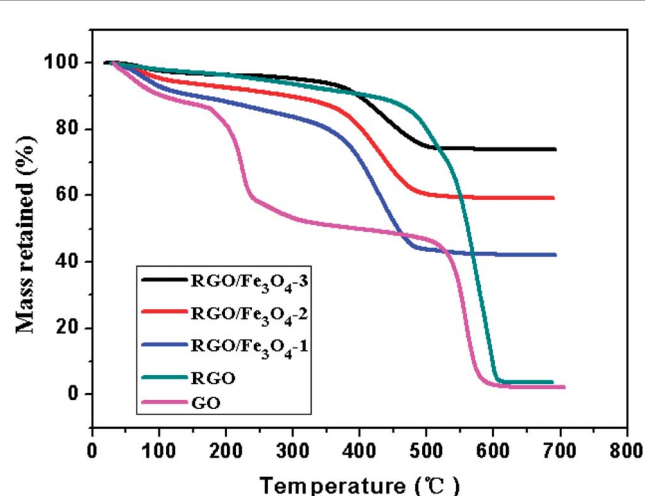


Fig. 6 TG curves of GO, RGO and RGO-Fe₃O₄ composites in air atmosphere.

The field dependence of magnetization for the Fe₃O₄ NPs and RGO–Fe₃O₄ composites was measured by a vibrating sample magnetometer (VSM) at room temperature, as shown in Fig. 7. The magnetic parameters corresponding to Fig. 7 are also shown in Table 1, including the saturation magnetization (M_s), coercivity (H_c), and remnant magnetization (M_r). It clearly shows that the M_s and M_r of the composites are increased while the H_c decreases with the increase of the Fe₃O₄ loading amount. The magnetization hysteresis loops of the Fe₃O₄ NPs and RGO–Fe₃O₄ composites are S-like, which is typical of a superparamagnetic material because of their high M_s and M_r , and low H_c ,³⁰ implying that there is no remaining magnetization when the applied magnetic field is removed.⁴⁴ The superparamagnetic properties of the composites further confirm that the oxide in the present investigation is Fe₃O₄ rather than γ -Fe₂O₃ and is supported by the XPS spectra for Fe 2p.³⁸ These composites could be separated from their dispersion within 15 seconds by holding the samples close to a magnet as shown in Fig. S3,[†] indicating that it is possible to manipulate these composites by an external magnetic field. Additionally, the M_s value of all the RGO–Fe₃O₄ composites is lower than that of the pure bulk Fe₃O₄, which can be attributed to the nanoscale size of the Fe₃O₄ NPs and the presence of RGO.^{23,44}

The electromagnetic parameters (relative complex permittivity, $\epsilon_r = \epsilon' - j\epsilon''$, and relative complex permeability, $\mu_r = \mu' - j\mu''$) of the wax composites containing 50 wt% of the samples were measured at room temperature for the investigation of microwave absorption properties of the Fe₃O₄ NPs and RGO–Fe₃O₄ composites, and are shown in Fig. 8a–d. The real permittivity (ϵ') and real permeability (μ') symbolize the storage ability of electric and magnetic energy, while the imaginary permittivity (ϵ'') and imaginary permeability (μ'') are related to the dissipation of electric and magnetic energy.²⁰

As shown in Fig. 8a, for the RGO–Fe₃O₄-1 composite, RGO–Fe₃O₄-2 composite and RGO–Fe₃O₄-3 composite, the values of ϵ' are in the range of 15.37–8.22, 12.30–7.89, and 11.30–7.89, respectively, in the frequency of 2–18 GHz, and are higher than the pure Fe₃O₄ NPs (in the range of 5.04–3.13). Meanwhile, for

Table 1 Magnetic properties of Fe₃O₄ NPs and RGO–Fe₃O₄ composites

Sample	M_s (emu g ⁻¹)	H_c (Oe)	M_r (emu g ⁻¹)
RGO–Fe ₃ O ₄ -1	27.6	114.9	4.8
RGO–Fe ₃ O ₄ -2	42.4	108.9	7.4
RGO–Fe ₃ O ₄ -3	54.3	103.5	9.0
Fe ₃ O ₄	80.2	42.9	9.5

the RGO–Fe₃O₄-1 composite, RGO–Fe₃O₄-2 composite and RGO–Fe₃O₄-3 composite, the values of ϵ'' are in the range of 7.67–0.88, 4.61–0.49, and 4.20–0.09, respectively, and are also higher than the pure Fe₃O₄ NPs (in the range of 1.21–0.10), as presented in Fig. 8b. It is observed that the samples with higher (RGO)/(Fe₃O₄) ratios show higher values of ϵ' and ϵ'' , which is related to higher conductivities. This is because more RGO sheets increase the electric polarization and electric conductivity, since ϵ_r is an expression of the polarizability of a material, which consists of dipolar polarization and electric polarization at microwave frequencies.¹

From Fig. 8c, for the RGO–Fe₃O₄-1 composite, RGO–Fe₃O₄-2 composite and RGO–Fe₃O₄-3 composite, we can see that the values of μ' are in the range of 1.18–0.83, 1.13–0.77, and 1.15–0.73, respectively, and thus have a very serious fluctuation in the 2–18 GHz frequency range. In addition, the values of μ'' for the RGO–Fe₃O₄-1 composite, RGO–Fe₃O₄-2 composite and RGO–Fe₃O₄-3 composite, as presented in Fig. 8d, are in the range of 0.87–0.03, 0.87–0.01, and 0.80 to (–0.13) in the frequency of 2–18 GHz, which are almost equal to the pure Fe₃O₄ NPs. Meanwhile, the μ'' for the RGO–Fe₃O₄ composites exhibits a broad resonance peak at 12.4–15.6 GHz with a maximum value at 14.8 GHz.

We have also calculated both the dielectric tangent loss ($\tan \delta_E = \epsilon''/\epsilon'$) and the magnetic tangent loss ($\tan \delta_M = \mu''/\mu'$) of the Fe₃O₄ NPs and RGO–Fe₃O₄ composites based on the permeability and permittivity of samples measured as above, shown in Fig. 8e and f. The values of $\tan \delta_E$ are mostly larger than 0.2 between 2 to 18 GHz, indicating that the dielectric loss occurs over the whole frequency range. These results suggest the RGO–Fe₃O₄ composites have distinct dielectric loss properties.

So far, there is no evidence to support the C–O–Fe bond affect on the microwave absorption properties. But the introduction of Fe₃O₄ NPs significantly enhanced the microwave absorption performance of the composites. In terms of electromagnetic theory, the dielectric loss of the RGO–Fe₃O₄ composites may be attributed to natural resonance, Debye dipolar relaxation, electron polarization relaxation and the unique layered nanostructures *etc.*^{1,5,45} First, the presence of the Fe₃O₄ NPs and residual defects and groups in RGO (discussed in the Raman analysis) does not significantly alter the graphene lamellar structure, but they can act as polarized centers, which are in favor of the electromagnetic energy absorption.¹ Second, the huge aspect ratio, layered structure and high conductivity of the RGO sheets are another reason why the RGO–Fe₃O₄ composites exhibit better absorbing abilities.¹ Thirdly, there are dipoles present in Fe₃O₄ especially when their sizes are at the

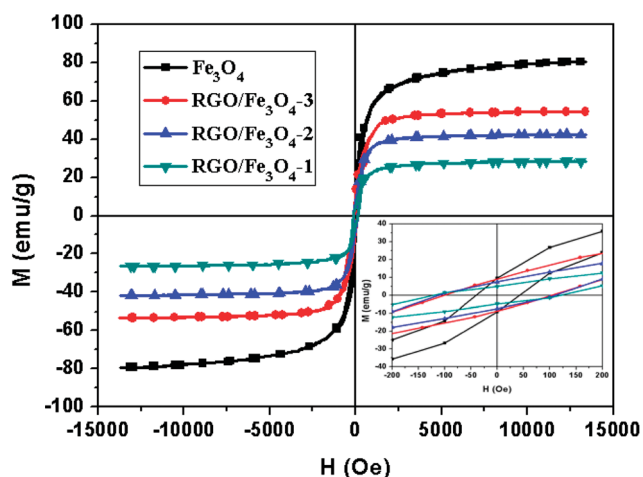


Fig. 7 Magnetization curves of Fe₃O₄ NPs and RGO–Fe₃O₄ composites at room temperature. Inset: the expanded low field magnetization curves.

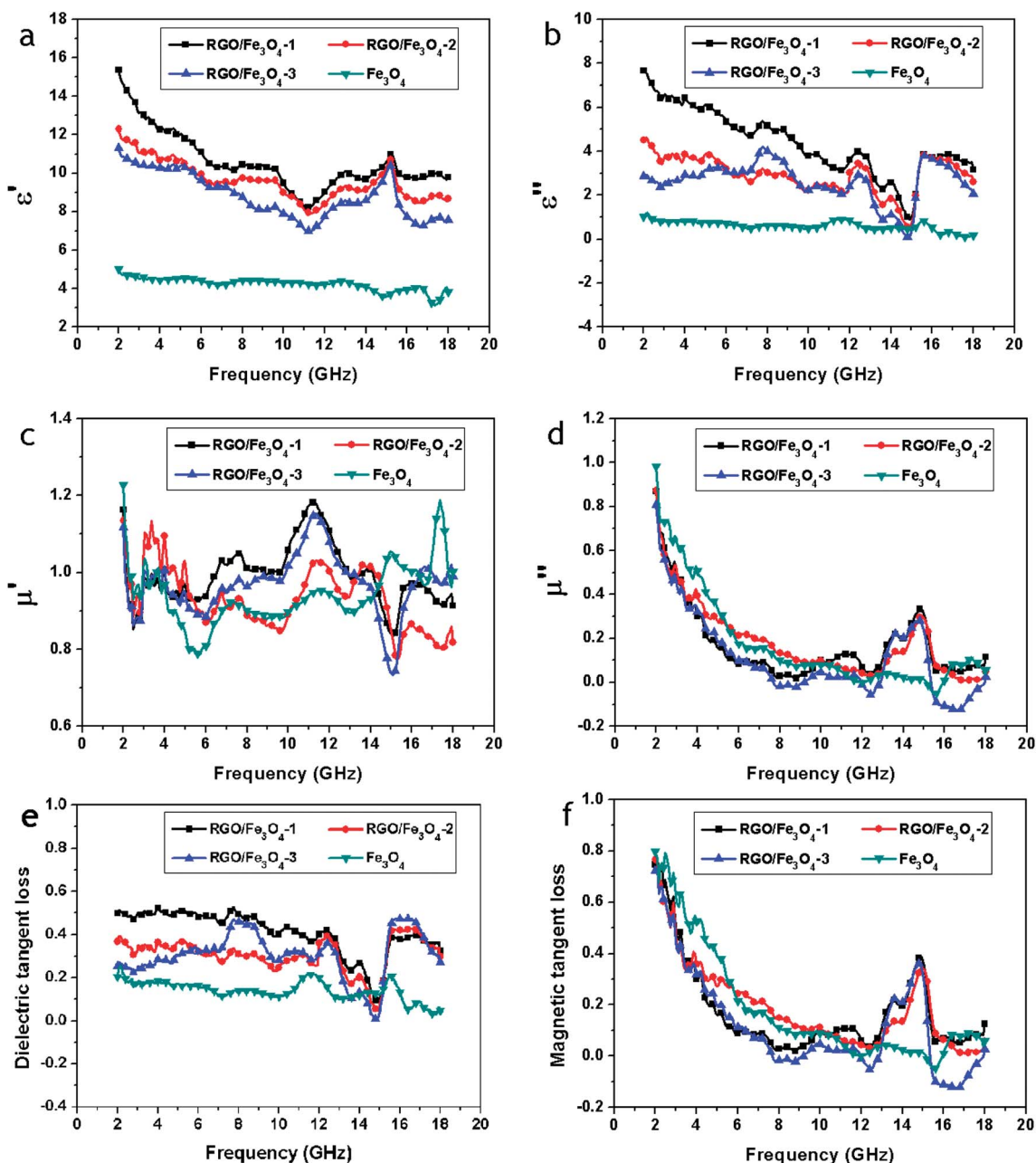


Fig. 8 Frequency dependence on real parts (a) and imaginary parts (b) of the complex permittivity, real parts (c) and imaginary parts (d) of the complex permeability, and the corresponding dielectric loss tangents (e) and magnetic loss tangents (f) of Fe_3O_4 and RGO- Fe_3O_4 composites.

nanoscale. The number of surface atoms with unsaturated bonds greatly increases as the size decreases, causing an increase of the dipoles. Consequently, the dipole polarizations can contribute to the dielectric loss.⁴ If we just make a mixture of GO with Fe_3O_4 followed by a reduction, the Fe_3O_4 NPs will conglomerate, then the dipole polarization will weaken. Moreover, the interfaces between the RGO lamella and Fe_3O_4 NPs are clearly observed, as shown in Fig. 1. The interfacial polarization and the associated relaxation contribute to the dielectric loss.¹ This is another reason why the RGO- Fe_3O_4 composites have stronger EM absorption than Fe_3O_4 NPs. Furthermore, according to the free electron theory, $\epsilon'' = 1/2\epsilon_0\pi\rho f$, where ϵ_0 is the

permittivity of a vacuum, ρ is the resistivity, f is the frequency of the microwave.^{30,46,47} The conductivity of graphene is high and a conducting network is formed, which enables a reduction in resistivity ρ of the composites and the ϵ'' becomes higher with the increasing content of graphene.³⁰ The resistivity of the nanocomposite will decrease due to the increase of the (RGO)/(Fe_3O_4) ratios and results in an increase in dielectric loss.⁴⁶ In addition, such higher dielectric loss is also related to electron hopping between Fe^{2+} and Fe^{3+} ions in the B sites as the EM wave field is applied, which has been reported previously.^{4,46}

We adopt the Debye dielectric relaxation model (Cole-Cole model) to further interpret the mechanisms of the permittivity

dispersion. According to its expression, the relative complex permittivity can be expressed by the following equation,^{1,5,47}

$$\epsilon_r = \epsilon_\infty + \frac{\epsilon_s - \epsilon_\infty}{1 + j2\pi f\tau} = \epsilon' - j\epsilon'' \quad (4)$$

where f , ϵ_s , ϵ_∞ , and τ are frequency, static permittivity, relative dielectric permittivity at the high-frequency limit, and polarization relaxation time, respectively. Thus, ϵ' and ϵ'' can be described by

$$\epsilon' = \epsilon_\infty + \frac{\epsilon_s - \epsilon_\infty}{1 + (2\pi f)^2\tau^2} \quad (5)$$

$$\epsilon'' = \frac{2\pi f\tau(\epsilon_s - \epsilon_\infty)}{1 + (2\pi f)^2\tau^2} \quad (6)$$

According to eqn (2) and (3), the relationship between ϵ' and ϵ'' can be deduced,

$$\left(\epsilon' - \frac{\epsilon_s + \epsilon_\infty}{2}\right)^2 + (\epsilon'')^2 = \left(\frac{\epsilon_s - \epsilon_\infty}{2}\right)^2 \quad (7)$$

Thus, the plot of ϵ' versus ϵ'' would be a single semicircle, generally denoted as the Cole–Cole semicircle. Each semicircle corresponds to one Debye relaxation process. Fig. 9a–c show the ϵ' – ϵ'' curves of the RGO–Fe₃O₄ composites.

At least four Cole–Cole semicircles are clearly found in the ϵ' – ϵ'' curves of the RGO–Fe₃O₄ composites. This suggests that

there are multiple relaxation processes for the RGO–Fe₃O₄ composites, representing the contribution of the Debye relaxation process to the enhanced dielectric properties of the RGO–Fe₃O₄ composites.¹ In the composites, the existence of interfaces gives rise to interfacial polarization or the Maxwell–Wagner effect. This phenomenon appears in heterogeneous media due to the accumulation of charges at the interfaces and the formation of large dipoles on particles or clusters.⁴⁷ Evidently, the Cole–Cole semicircle is distorted, suggesting that except for the dielectric relaxation, other mechanisms such as conductance loss, interfacial polarization and oxygen defects may contribute to the permittivity spectra.⁴⁵

It reveals that the value of the magnetic loss ($\tan \delta_M$) of the RGO–Fe₃O₄ composites first declines at 2.0–8.0 GHz, remains in the 8.0–12.4 GHz range, and then exhibits broad resonance peaks at 12.4–15.6 GHz with a maximum value at 14.8 GHz, as presented in Fig. 8f. The magnetic loss is related to the eddy current effect and the anisotropy energy of the composites. For the ferromagnetic absorber, the microwave absorption properties are usually subject to degradation caused by the eddy current effect in the high-frequency region. The μ'' can be expressed by⁴

$$\mu'' \approx 2\pi\mu_0(\mu')^2\sigma d^2f/3 \quad (8)$$

where σ and μ_0 are the electric conductivity and the permeability in vacuum, respectively. Thus, C_0 can be described by

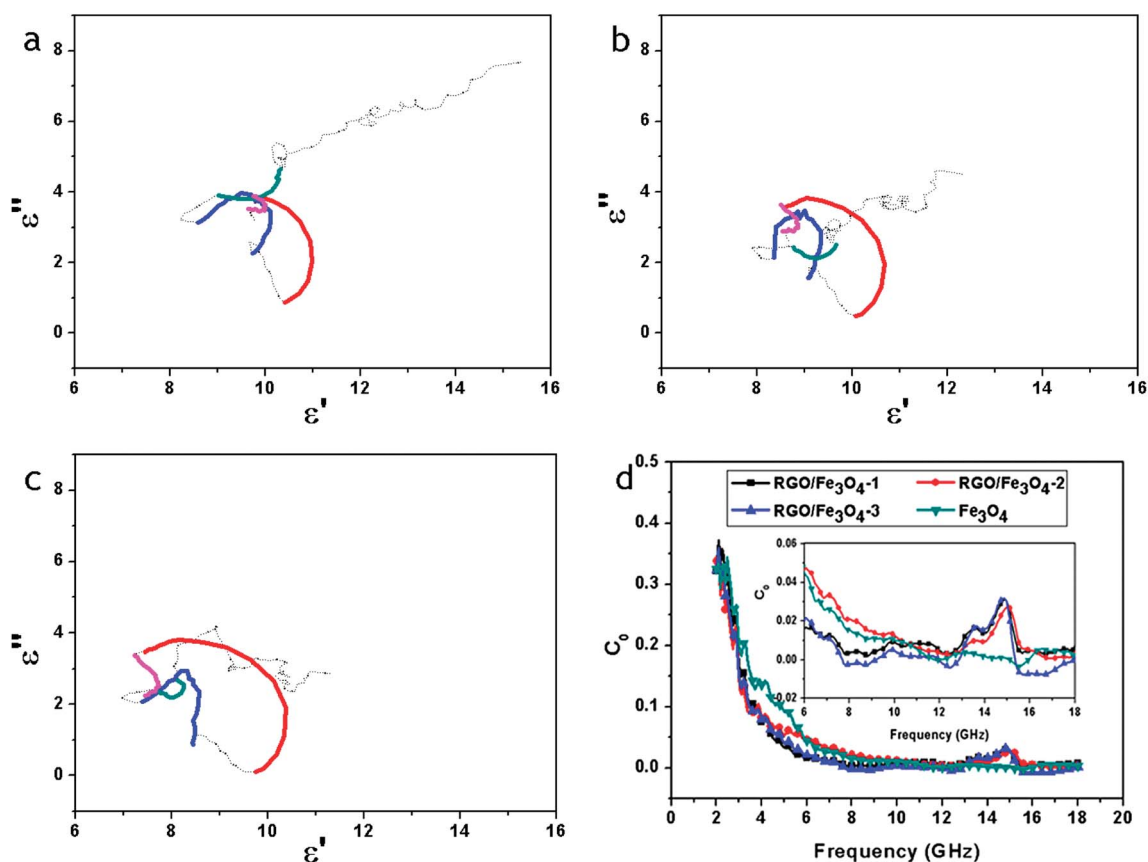


Fig. 9 ϵ' – ϵ'' curves of RGO–Fe₃O₄-1 composite (a), RGO–Fe₃O₄-2 composite (b) and RGO–Fe₃O₄-3 composite (c); C_0 – f curves of Fe₃O₄ NPs and RGO–Fe₃O₄ composites (d), inset is the expansion of the curves.



Fig. 10 The calculated reflection losses for RGO-Fe₃O₄-1 (a), RGO-Fe₃O₄-2 (b) and RGO-Fe₃O₄-3 (c) with different thicknesses in the frequency range of 2–18 GHz.

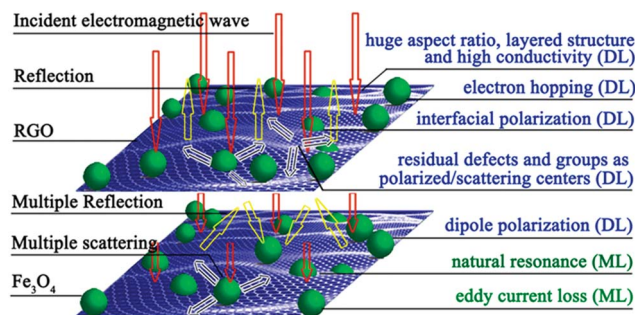


Fig. 11 Schematic of possible microwave absorbing mechanisms of the RGO-Fe₃O₄ composites.

$$C_0 = \mu''(\mu')^{-2}f^{-1} = 2\pi\mu_0\sigma d^2/3 \quad (9)$$

If the magnetic loss results from eddy current loss, the values of C_0 are constant when the frequency is varied.⁴⁸

Fig. 9d shows the C_0 - f curves of the pure Fe₃O₄ NPs and the RGO-Fe₃O₄ composites. For the pure Fe₃O₄ NPs, the values of C_0 are almost constant in the high-frequency region (8–18 GHz). However, for the RGO-Fe₃O₄ composites, those values of C_0 exhibit a broad peak in the 12.4–15.6 GHz frequency range. This implies that the pure Fe₃O₄ NPs have a significant eddy current effect in the high-frequency region (8–18 GHz), and the range of eddy current effect will be decreased with the introduction of RGO. Besides, the magnetic loss in the RGO-Fe₃O₄ composites was mainly caused by the natural resonance in the frequency range 2.0–8.0 GHz and 12.4–15.6 GHz and the eddy current effect in the frequency range 8.0–12.4 GHz and 15.6–18.0 GHz.

The natural resonance in the frequency range 2.0–8.0 GHz and 12.4–15.6 GHz can be attributed to the small size effect and the introduction of RGO. According to the natural-resonance equation,^{4,49} for ferromagnetic materials,

$$2\pi f_r = rH_a \quad (10)$$

$$H_a = 4|K_1|/3\mu_0M_s \quad (11)$$

where r is the gyromagnetic ratio, H_a is the anisotropy energy, and $|K_1|$ is the anisotropy coefficient. On one hand, the anisotropy energy of small materials, especially on the nano-scale, would remarkably increase due to the surface anisotropic field by the small size effect.^{49,50} On the other hand, the M_s value

of the RGO-Fe₃O₄ composites is much lower than that of the pure Fe₃O₄ NPs (Fig. 7), thus the anisotropy energy of the RGO-Fe₃O₄ composites is higher than that of the pure Fe₃O₄ NPs. The higher anisotropy energy is helpful for the improvement of microwave absorption properties.⁴

To further study the microwave absorption properties, the reflection loss (R_L) was calculated by the following equations based on the transmit-line theory:

$$Z_{in} = \sqrt{\frac{\mu_r}{\epsilon_r}} \tanh \left[j \left(\frac{2\pi f d}{c} \right) \sqrt{\epsilon_r \mu_r} \right] \quad (12)$$

$$R_L(\text{dB}) = 20 \log \left| \frac{Z_{in} - 1}{Z_{in} + 1} \right| \quad (13)$$

where Z_{in} is the input impedance of the absorber, μ_r and ϵ_r are respectively the relative complex permeability and permittivity, f is the frequency of microwaves, d is the thickness of the absorber, and c is the velocity of electromagnetic waves in free space. The impedance matching condition is determined by the combination of the six parameters: ϵ' , ϵ'' , μ' , μ'' , f , and d .²⁰

Fig. 10 shows the calculated theoretical reflection loss of the RGO-Fe₃O₄ composites with different thickness in the range of 2–18 GHz. It can be observed that the thickness of the absorbers has a great influence on the microwave absorbing properties. In addition, the maximum R_L absorption gradually appeared to shift toward a lower frequency with the increase of thickness.¹

As shown in Fig. 10a, the maximum R_L reaches -38.4 dB at 11.6 GHz for the RGO-Fe₃O₄-1 composite with a thickness of only 2.1 mm. With a thickness of 1.5 mm, the maximum R_L which can be achieved is -24.2 dB at 16.6 GHz, and a bandwidth of R_L less than -10 dB (90% absorption¹⁹) can reach up to 4.2 GHz (from 13.8 to 18.0 GHz). The absorption bandwidth with R_L below -10 dB is up to 12.7 GHz (from 5.3 to 18.0 GHz) with a thickness in the range of 1.5–4.0 mm. For the RGO-Fe₃O₄-2 composite (Fig. 10b), the maximum R_L reaches -34.9 dB at 9.4 GHz with a thickness of only 2.8 mm, and a bandwidth of R_L less than -10 dB can reach up to 3.2 GHz (from 8.0 to 11.2 GHz). Whereas for the RGO-Fe₃O₄-3 composite (Fig. 10c), the maximum R_L even reaches -44.6 dB at 6.6 GHz with a thickness of 3.9 mm, and a bandwidth of R_L less than -10 dB can reach up to 2.0 GHz (from 5.7 to 7.7 GHz). With a thickness of 2.0 mm, the maximum R_L can be achieved at -15.9 dB at 13.2 GHz, and a bandwidth of R_L less than -10 dB

can reach up to 4.3 GHz (from 12.2 to 16.5 GHz). It is observed that for the RGO-Fe₃O₄ composites, the optimal thickness will increase whereas the matching frequency will decrease when the (RGO)/(Fe₃O₄) ratio decreases. As previously stated, we conclude that the as-prepared RGO-Fe₃O₄ composites show an excellent absorption performance in the range of 2–18 GHz and the maximum R_L can be adjusted to a different frequency by controlling the thickness. Our RGO-Fe₃O₄ composites exhibit excellent microwave absorption performance in comparison with pure RGO,^{3,30} Fe₃O₄ NPs (Fig. S4†) and RGO composites reported.¹²

The excellent microwave absorbing performance of the RGO-Fe₃O₄ composites is mainly attributed to two key factors: impedance matching and electromagnetic wave attenuation.³⁰ On one hand, the introduction of Fe₃O₄ NPs has lowered the ϵ_r of the RGO, and improved the equality of the ϵ_r and μ_r , which helps to improve the level of impedance matching. On the other hand, the RGO-Fe₃O₄ composites have strong electromagnetic wave attenuation, which is determined by their dielectric loss (DL) and magnetic loss (ML) (discussed in Fig. 8e and f). Besides, the enormous aspect ratio, layered structure, and the existence of residual defects and groups of the RGO-Fe₃O₄ composites could cause multiple reflections, which will further enhance the microwave absorbing ability of the composites.^{30,41} In general, the enhanced microwave absorbing performance of the composites is attributed to the compensatory effect of the graphene and Fe₃O₄ NPs.³⁰ To further give a visual demonstration of the microwave absorbing mechanism as discussed above, a schematic diagram is presented in Fig. 11. From the above, the results demonstrate that the RGO-Fe₃O₄ composites could be used as a microwave absorbing material.

4. Conclusion

In summary, the RGO-Fe₃O₄ composites with an obviously enhanced microwave absorption property have been successfully synthesized by a rational one-pot simplified co-precipitation route, which avoided the usage of an inert gas and any additional chemical agents (such as surfactants). The results show that the Fe₃O₄ NPs are very well dispersed on the surface of the RGO nanosheets. The RGO-Fe₃O₄ composites remarkably improved the electromagnetic performance in comparison with the pure RGO, Fe₃O₄ NPs and RGO composites reported, which is attributed to effective complementarities between the dielectric loss and the magnetic loss. The dielectric loss of the composites can be attributed to natural resonance, Debye dipolar relaxation, electron polarization relaxation, interfacial polarization and the unique layered nanostructures *etc.* The magnetic loss of the RGO-Fe₃O₄ composites is mainly caused by the natural resonance in the frequency range 2.0–8.0 GHz and 12.4–15.6 GHz, and the eddy current effect in the frequency range 8.0–12.4 GHz and 15.6–18.0 GHz. For the RGO-Fe₃O₄-3 composite, the maximum R_L reaches -44.6 dB at 6.6 GHz with a thickness of 3.9 mm, and a bandwidth of R_L less than -10 dB can reach up to 4.3 GHz (from 12.2 to 16.5 GHz) with a thickness of 2.0 mm. Additionally, the microwave absorption property can be tuned easily by varying the (RGO)/(Fe₃O₄) ratio and layer

thickness of the samples. It is believed that such composites will find wide applications in microwave absorbing area.

Acknowledgements

This work was supported by the Spaceflight Foundation of China (no. 2011XW110001C110001), the Spaceflight Innovation Foundation of China (no. 2011XT110002C110002) and the Basic Research Foundation of Northwestern Polytechnical University (no. JC201269).

References

- 1 D. Z. Chen, G. S. Wang, S. He, J. Liu, L. Guo and M. S. Cao, *J. Mater. Chem. A*, 2013, **1**, 5996–6003.
- 2 J. W. Liu, R. C. Che, H. J. Chen, F. Zhang, F. Xia, Q. S. Wu and M. Wang, *Small*, 2012, **8**, 1214–1221.
- 3 C. Wang, X. J. Han, P. Xu, X. L. Zhang, Y. C. Du, S. R. Hu, J. Y. Wang and X. H. Wang, *Appl. Phys. Lett.*, 2011, **98**, 072906.
- 4 C. L. Zhu, M. L. Zhang, Y. J. Qiao, G. Xiao, F. Zhang and Y. J. Chen, *J. Phys. Chem. C*, 2010, **114**, 16229–16235.
- 5 H. L. Yu, T. S. Wang, B. Wen, M. M. Lu, Z. Xu, C. L. Zhu, Y. J. Chen, X. Y. Xue, C. W. Sun and M. S. Cao, *J. Mater. Chem.*, 2012, **22**, 21679–21685.
- 6 X. A. Li, B. Zhang, C. H. Ju, X. J. Han, Y. C. Du and P. Xu, *J. Phys. Chem. C*, 2011, **115**, 12350–12357.
- 7 Y. C. Qing, W. C. Zhou, F. Luo and D. M. Zhu, *Carbon*, 2010, **48**, 4074–4080.
- 8 X. Bai, Y. H. Zhai and Y. Zhang, *J. Phys. Chem. C*, 2011, **115**, 11673–11677.
- 9 L. Wang, F. He and Y. Z. Wan, *J. Alloys Compd.*, 2011, **509**, 4726–4730.
- 10 D. Y. Pan, J. C. Zhang, Z. Li and M. H. Wu, *Adv. Mater.*, 2010, **22**, 734–738.
- 11 H. X. Chang, G. F. Wang, A. Yang, X. M. Tao, X. Q. Liu, Y. D. Shen and Z. J. Zheng, *Adv. Funct. Mater.*, 2010, **20**, 2893–2902.
- 12 E. L. Ma, J. J. Li, N. Q. Zhao, E. Z. Liu, C. N. He and C. S. Shi, *Mater. Lett.*, 2013, **91**, 209–212.
- 13 G. Goncalves, P. A. A. P. Marques, C. M. Granadeiro, H. I. S. Nogueira, M. K. Singh and J. Grácio, *Chem. Mater.*, 2009, **21**, 4796–4802.
- 14 S. Liu, J. Q. Tian, L. Wang, H. L. Li, Y. W. Zhang and X. P. Sun, *Macromolecules*, 2010, **43**, 10078–10083.
- 15 G. P. Kim, I. Nam, N. D. Kim, J. Park, S. Park and J. Yi, *Electrochem. Commun.*, 2012, **22**, 93–96.
- 16 A. Prakash, S. Chandra and D. Bahadur, *Carbon*, 2012, **50**, 4209–4219.
- 17 Y. Q. Zhan, F. B. Meng, Y. J. Lei, R. Zhao, J. C. Zhong and X. B. Liu, *Mater. Lett.*, 2011, **65**, 1737–1740.
- 18 Y. Q. Zhan, R. Zhao, Y. J. Lei, F. B. Meng, J. C. Zhong and X. B. Liu, *Appl. Surf. Sci.*, 2011, **257**, 4524–4528.
- 19 Y. Yang, X. L. Liu, Y. Yang, W. Xiao, Z. W. Li, D. S. Xue, F. S. Li and J. Ding, *J. Mater. Chem. C*, 2013, **1**, 2875–2885.
- 20 W. C. Zhou, X. J. Hu, X. X. Bai, S. Y. Zhou, C. Y. Sun, J. Yan and P. Chen, *ACS Appl. Mater. Interfaces*, 2011, **3**, 3839–3845.

- 21 S. H. Liu, R. M. Xing, F. Lu, R. K. Rana and J. J. Zhu, *J. Phys. Chem. C*, 2009, **113**, 21042–21047.
- 22 J. S. Chen, Y. M. Zhang and D. Lou, *ACS Appl. Mater. Interfaces*, 2011, **3**, 3276–3279.
- 23 J. Su, M. H. Cao, L. Ren and C. W. Hu, *J. Phys. Chem. C*, 2011, **115**, 14469–14477.
- 24 W. Hummers, Jr and R. E. Offeman, *J. Am. Chem. Soc.*, 1958, **80**, 1339.
- 25 M. Z. Kassaei, E. Motamedi and M. Majidi, *Biochem. Eng. J.*, 2011, **172**, 540–549.
- 26 P. S. Teo, H. N. Lim, N. M. Huang, C. H. Chia and I. Harrison, *Ceram. Int.*, 2012, **38**, 6411–6416.
- 27 X. P. Shen, J. L. Wu, S. Bai and H. Zhou, *J. Alloys Compd.*, 2010, **506**, 136–140.
- 28 Y. C. Si and E. T. Samulski, *Chem. Mater.*, 2008, **20**, 6792–6797.
- 29 H. Wei, W. s. Yang, Q. Xi and X. Chen, *Mater. Lett.*, 2012, **82**, 224–226.
- 30 X. Sun, J. P. He, G. X. Li, J. Tang, T. Wang, Y. X. Guo and H. R. Xue, *J. Mater. Chem. C*, 2013, **1**, 765–777.
- 31 G. B. Sun, B. X. Dong, M. H. Cao, B. Q. Wei and C. W. Hu, *Chem. Mater.*, 2011, **23**, 1587–1593.
- 32 Y. Q. Zhan, F. B. Meng, X. L. Yang, R. Zhao and X. B. Liu, *Mater. Sci. Eng., B*, 2011, **176**, 1333–1339.
- 33 P. B. Liu, Y. Huang and L. Wang, *Mater. Lett.*, 2013, **91**, 125–128.
- 34 H. L. Guo, X. F. Wang, Q. Y. Qian, F. B. Wang and X. H. Xia, *ACS Nano*, 2009, **3**, 2653–2659.
- 35 H. K. Jeong, Y. P. Lee, R. J. W. E. Lahaye, M. H. Park, K. H. An, I. J. Kim, C. W. Yang, C. Y. Park, R. S. Ruoff and Y. H. Lee, *J. Am. Chem. Soc.*, 2008, **130**, 1362–1366.
- 36 H. P. Cong, X. C. Ren, P. Wang and S. H. Yu, *ACS Nano*, 2012, **6**, 2693–2703.
- 37 C. Z. Zhu, S. J. Guo, Y. X. Fang and S. J. Dong, *ACS Nano*, 2010, **4**, 2429–2437.
- 38 L. L. Ren, S. Huang, W. Fan and T. X. Liu, *Appl. Surf. Sci.*, 2011, **258**, 1132–1138.
- 39 Y. Z. Xue, B. Wu, L. Jiang, Y. L. Guo, L. P. Huang, J. J. Chen, J. H. Tan, D. C. Geng, B. R. Luo, W. P. Hu, G. Yu and Y. Q. Liu, *J. Am. Chem. Soc.*, 2012, **134**, 11060–11063.
- 40 F. J. Zhang, J. Liu, K. Zhang, W. Zhao, W. K. Jang and W. C. Oh, *Korean J. Chem. Eng.*, 2012, **29**, 989–993.
- 41 G. M. Rutter, J. N. Crain, N. P. Guisinger, T. Li, P. N. First and J. A. Stroschio, *Science*, 2007, **317**, 219–222.
- 42 X. B. Fan, W. C. Peng, Y. Li, X. Y. Li, S. L. Wang, G. L. Zhang and F. B. Zhang, *Adv. Mater.*, 2008, **20**, 4490–4493.
- 43 D. Zhou, T. L. Zhang and B. H. Han, *Microporous Mesoporous Mater.*, 2013, **165**, 234–239.
- 44 F. He, J. T. Fan, D. Ma, L. M. Zhang, C. Leung and H. L. Chan, *Carbon*, 2010, **48**, 3139–3144.
- 45 H. J. Wu, L. D. Wang, S. L. Guo, Y. M. Wang and Z. Y. Shen, *Mater. Chem. Phys.*, 2012, **133**, 965–970.
- 46 Y. J. Chen, G. Xiao, T. S. Wang, Q. Y. Ouyang, L. H. Qi, Y. Ma, P. Gao, C. L. Zhu, M. S. Cao and H. B. Jin, *J. Phys. Chem. C*, 2011, **115**, 13603–13608.
- 47 S. He, G. S. Wang, C. Lu, J. Liu, B. Wen, H. Liu, L. Guo and M. S. Cao, *J. Mater. Chem. A*, 2013, **1**, 4685–4692.
- 48 X. F. Zhang, X. L. Dong, H. Huang, Y. Y. Liu, W. N. Wang, X. G. Zhu, B. Lv and J. P. Lei, *Appl. Phys. Lett.*, 2006, **89**, 053115.
- 49 Y. J. Chen, P. Gao, R. X. Wang, C. L. Zhu, L. J. Wang, M. J. Cao and H. B. Jin, *J. Phys. Chem. C*, 2009, **113**, 10061–10064.
- 50 D. L. Leslie-Pelecky and R. D. Rieke, *Chem. Mater.*, 1996, **8**, 1770–1783.