

Journal of Materials Chemistry A

Materials for energy and sustainability

Accepted Manuscript

This article can be cited before page numbers have been issued, to do this please use: A. Trifoglio, M. Melucci, A. Pintus, S. Khaliha, F. Blancato, M. Agnes, L. Aversa, R. Verucchi, T. D. Marforio, N. Burduja, G. Nocito, A. Mazzaglia and M. Calvaresi, *J. Mater. Chem. A*, 2025, DOI: 10.1039/D5TA08410E.



This is an Accepted Manuscript, which has been through the Royal Society of Chemistry peer review process and has been accepted for publication.

Accepted Manuscripts are published online shortly after acceptance, before technical editing, formatting and proof reading. Using this free service, authors can make their results available to the community, in citable form, before we publish the edited article. We will replace this Accepted Manuscript with the edited and formatted Advance Article as soon as it is available.

You can find more information about Accepted Manuscripts in the [Information for Authors](#).

Please note that technical editing may introduce minor changes to the text and/or graphics, which may alter content. The journal's standard [Terms & Conditions](#) and the [Ethical guidelines](#) still apply. In no event shall the Royal Society of Chemistry be held responsible for any errors or omissions in this Accepted Manuscript or any consequences arising from the use of any information it contains.

ARTICLE

Received 00th January 20xx,

Poly-Cyclodextrin modified graphene oxide for PFAS removal from drinking water

Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

Andrea Trifoglio,^{a,e,*} Angela Pintus,^{a,f} Sara Khaliha,^a Francesco Blancato,^a Marco Agnes,^a Lucrezia Aversa,^b Roberto Verucchi,^b Tainah Dorina Marforio,^{e,g} Nina Burduja,^{c,d} Giuseppe Nocito,^c Antonino Mazzaglia,^c Matteo Calvaresi,^{e,g} Manuela Melucci^{a*}

Per- and polyfluoroalkyl substances (PFAS) are persistent organic pollutants of growing concern due to their widespread occurrence in drinking water and resistance to conventional remediation technologies. Granular activated carbon (GAC), the current benchmark adsorbent, exhibits limited efficiency toward short- and medium-chain PFAS and slow adsorption kinetics. Here, we report the synthesis of a graphene oxide (GO) material covalently grafted with an amino-functionalized β -cyclodextrin polymer crosslinked with epichlorohydrin (GO–Poly– β CD). The hybrid material was characterized and evaluated for the adsorption of a mixture of PFAS with varying chain lengths (CF)_{3–9} and functional groups under environmentally relevant conditions. GO–Poly– β CD outperformed pristine GO, showing enhanced affinity for medium-chain PFAS and removal efficiencies up to 97% for PFHpA (CF)₆ and PFHxS (CF)₆ after only 15 min of contact. At neutral pH, PFPeA (CF)₄ removal reached 44%, compared to no removal by GO, while under acidic conditions its removal increased to 80%. Notably, PFBA (CF)₃, which was not removed at neutral pH, exhibited 31% removal under acidic conditions. Molecular dynamics simulations revealed a cooperative adsorption mechanism in which PFAS molecules are stabilized through pH-responsive conformational rearrangements that strengthen interactions between the hybrid material and the contaminants, primarily via van der Waals and hydrophobic forces. Compared to GAC, GO–Poly– β CD achieved a higher overall PFAS uptake (4.0 $\mu\text{g g}^{-1}$ vs. 1.3 $\mu\text{g g}^{-1}$ for GAC) within the same contact time, demonstrating its potential for rapid and efficient PFAS removal in real drinking water treatment applications.

Introduction

Per- and Poly-fluoro alkyl substances (PFAS) have attracted global concern due to their widespread use in everyday products, their environmental and biological persistence and widely reported effects on human health and ecosystems.^{1–3} PFAS are a class of synthetic organo-fluorine compounds characterized by the presence of fully or partially fluorinated alkyl-chains with a carboxylate or sulphonate ending group.⁴ PFAS are widely used as additive in food packaging, coatings, non-stick cookware, waterproof clothing and firefighting foams.⁵ The disposal of everyday products inevitably results in the release of PFAS into various environmental compartments and, ultimately, into drinking water sources through multiple

pathways.¹ Being the C-F chemical bond very stable, their persistency in the environment is exceptional making PFAS the so called “forever chemicals”.⁶ The Environmental Working Group reported that in 2018 over 16 million people in the United States were exposed to PFAS contaminated drinking water with concentrations exceeding 10 ngL^{–1}.³ The removal of PFAS from drinking water remains one of the key challenges in water purification. Among the treatment strategies currently implemented in municipal drinking water plants, adsorption is the most effective approach. In particular, granular activated carbon (GAC) represents the most widely employed adsorbent for PFAS mitigation.^{7–9} GAC is a non-specific adsorbent that exhibits moderate affinity toward long-chain PFAS but limited effectiveness for their short- and medium-chain counterparts.^{10, 11} Moreover, its adsorption process is characterized by slow, diffusion-limited kinetics, which represents a significant drawback under real plant flow conditions.^{12, 13} New technologies for GAC replacement or integration are currently matter of great interest. Among emerging materials, graphene oxide (GO) nanosheets have attracted considerable attention as promising adsorbents for the removal of emerging contaminants from drinking water, offering new opportunities for advanced treatment technologies.^{10, 14} The most relevant features of GO are its large specific surface area and the abundance of tunable oxygen-containing functional groups on

^a Institute for Organic Synthesis and Photoreactivity (ISOF-CNR), Bologna.^b Institute of Materials for electronics and Magnetism (IMEM-CNR), Trento c/o Fondazione Bruno Kessler.^c Institute for Nanostructured Materials (ISMN-CNR), Messina.^d Department of Chemical, Biological, Pharmaceutical and Environmental Sciences (ChiBioFarAm), University of Messina, Messina, Italy.^e Department of Chemistry ‘Giacomo Ciamician’, University of Bologna, Italy.^f Dipartimento di Physics, University of Modena e Reggio Emilia, Modena, Italy.^g IRCCS Azienda Ospedaliero – Universitaria di Bologna, Preclinical & Translational Research in Oncology Lab (PRO), Bologna, Italy.

* These authors contributed equally

† Supplementary Information available: see DOI: 10.1039/x0xx00000x



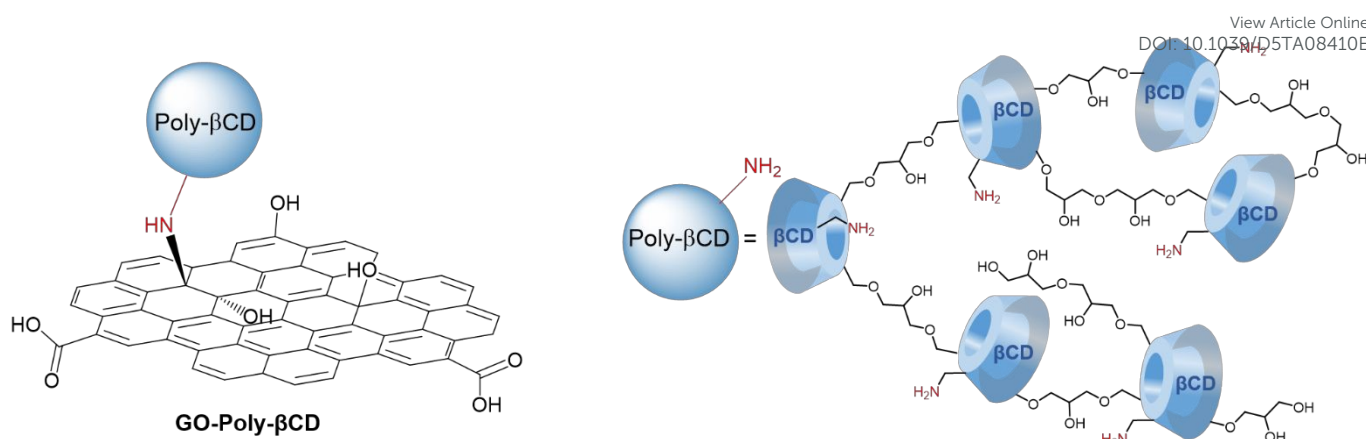


Figure 1. Chemical structure of GO-Poly-βCD and Poly-βCD.

the nanosheet surface, which enable versatile covalent functionalization.¹⁵

GO nanosheets have previously been explored for PFAS removal using a two-step approach involving batch adsorption followed by tandem microfiltration.¹⁰ Furthermore, some of us demonstrated that incorporating GO into polysulfone hollow fiber membranes (now commercially available as Graphisulfone^{®16}) provides a new integrated filter model combining adsorption and ultrafiltration mechanisms and allowing the removal of long-chain PFAS from drinking water.¹³ As a step forward, it has been demonstrated that the covalent functionalization of GO nanosheets enables the tuning of adsorption selectivity and significantly enhances the removal of contaminants that exhibit poor affinity for unmodified GO, including organic emerging pollutants with performances overcoming those achieved by GAC.^{10, 17, 18} For instance, functionalization of GO with N,N-dimethylethylenediamine (GO-DMEN),¹⁹ provided nanosheets with high affinity for medium- and long-chain PFAS. Studies based on theoretical modelling demonstrated that adsorption depends on multiple interactions including hydrophobic, electrostatic and van der Waals forces.¹³ In particular, the interaction between non-functionalized GO and PFAS is hindered by electrostatic repulsion between the negatively charged nanosheet and the negatively charged PFAS molecules.¹⁶ Hence, PFAS with longer hydrophobic chains are more efficiently adsorbed, whereas shorter, more hydrophilic PFAS are not removed.

Cyclodextrins are cyclic oligosaccharides made up by glucose subunits linked by α -1,4-glycosidic bond. They have a toroidal structure with hydrophobic inner cavity and a hydrophilic outer surface.²⁰ Cyclodextrins,^{21, 22} in particular β -cyclodextrins^{23, 24} (i.e. consisting in seven glucose subunits) can form inclusion complexes with PFAS through non-covalent forces such as van der Waals interactions and hydrophobic effects. Some of us, recently demonstrated that covalent modification of GO with an amino-derived β CD and a C6 aliphatic linker to GO nanosheets, enables the efficient removal of perfluorobutanoic acid (PFBA, (CF₃)₂₅ thank to the formation of a sandwich-like complex, where PFBA is stably confined between the GO nanosheet and the β CD cavity. Interestingly, β CD can serve as a monomer for the synthesis of β CD-based polymers through

various types of cross-linkers.²⁶ PFAS adsorption has been demonstrated for β CD polymers containing fluorinated aromatic cross-linkers,^{27, 28} which raise environmental concerns similar to those of the targeted PFAS.²⁹ Moreover, the partial water solubility of poly-cyclodextrins makes their direct application as sorbents under real conditions unrealistic.³⁰ To overcome this limitation, we propose here a novel polymeric material insoluble in water and based on a non-fluorinated chain (Poly- β CD from now on) and its use for PFAS removal from drinking water (Figure 1). GO-Poly- β CD adsorption mechanism was investigated by combining adsorption experiments under different pH conditions with molecular dynamics (MD) simulations.

Experimental section

Materials

GO was purchased by Layer One (Norway) and used without further purification. Before using, GO was sonicated in ultrapure water to exfoliate the bulk material into monolayer (>99%) with lateral size of few micrometers. GAC was purchased from CABOT Norit Spa (Ravenna, Italy, Norit GAC 830 AF, MB index min 240 mg/g, BET surface area >1000 m²/g) and used without further purification. PFAS standards were purchased by Agilent Technologies (Santa Clara, CA, US). The experiments on PFAS were carried out by using polypropylene vials. Poly- β CD was kindly provided Cyclolab (Budapest).

Synthesis of Poly- β -cyclodextrin (Poly- β CD)

Poly- β CD was provided as chlorohydrated salt of amino β -cyclodextrin polymer crosslinked with epichlorohydrin (average M_w = 25 kDa) and with 70 % of β -cyclodextrin content. The salt was treated with NaHCO₃ and dialysed for 72 hours in ultrapure water then lyophilized and used in the next synthetic step.

Synthesis and purification of GO-Poly- β CD

The synthesis of GO-Poly- β CD was carried out via epoxide ring-opening reaction.²⁵

Briefly, 100 mg of GO were dispersed in 30 mL of a 1:1 H₂O:EtOH solution by sonication for 2 hours to obtain a stable aqueous suspension. Separately, 200 mg of Poly- β CD-NH₂ were dissolved in 100 mL of distilled water and added to the GO suspension



(GO:Poly- β CD mass ratio 1:2, total volume 130 mL). The resulting mixture was refluxed at 80 °C for 12 hours. Upon completion of the reaction, the crude product was purified by centrifugation (9000 rpm, 15 minutes per cycle, eight cycles), and the final material was obtained by lyophilization.

Synthesis and purification of GO-Control

The synthesis of GO-Control was carried out under the same conditions as GO-Poly- β CD, but in the absence of Poly- β CD-NH₂. Briefly, 100 mg of GO were dispersed in 130 mL of a 1:1 H₂O:EtOH solution by sonication for 2 h to obtain a stable suspension. The mixture was then refluxed at 80 °C for 12 h. After completion of the reaction, the crude product was purified by centrifugation (9000 rpm, 15 min per cycle, eight cycles) and lyophilized.

Characterization

ATR-FTIR spectra were recorded with Agilent Cary 630 FTIR Spectrophotometer, and the spectra are expressed by wavenumber (cm⁻¹).

TGA were recorded with PerkinElmer Thermogravimetric Analyzer TGA 4000 by PerkinElmer, in air atmosphere, from 30 °C to 800 °C, with a scanning temperature of 10 °C/min.

Elemental analysis was performed on powder materials by using an Elementar Unicube Elemental analyser, method GRAPHITE.

XPS was performed using the Mg-K α emission at 1253.6 eV of a non-monochromatized X-ray photon source. Photoelectrons were analysed using a VSW HA100 electron energy analyser, resulting in a total energy resolution of 0.86 eV. Core level BEs were referred to the Au 4f_{7/2} core level signal (at 84.0 eV), obtained from a sputtered gold surface. The photoemission line shapes of selected core levels were analyzed using Voigt line shape deconvolution, after background subtraction of a Shirley function. The typical precision for energy peak position is ± 0.05 eV, while the uncertainty for FWHM is $< \pm 5\%$ and for area evaluation it is $\pm 5\%$. GO, GO-Control, and GO-Poly- β CD samples were mounted directly on the sample holder, while the Poly- β CD-NH₂ powder was supported on carbon tape.

Dynamic Light Scattering (DLS)

The mean hydrodynamic diameter (D_H), width of distribution (polydispersity index, PDI) and ζ -potential were measured with Malvern Zetasizer Nano ZS (Malvern Instruments, Malvern, UK), equipped with a He-Ne laser ($\lambda = 633$ nm, 4 mW). Measurements were performed using the non-invasive backscattering technique (NIBS) with an angle of 173 ° to the incident beam. For freshly prepared dispersions, 45 acquisitions at a temperature of 25 ± 1 °C for each measurement were collected. Deconvolution of the measured correlation curve (distribution of relaxation times and the cumulant analysis) into a dimensional average intensity distribution was obtained using the Laplace inversion algorithm.

Colloidal stability

The colloidal stability of GO-Poly- β CD was evaluated with an adapted turbidimetric method for a 250 μ g/mL dispersion in aqueous 1 mM NaCl. The sample was placed in a 1 cm quartz cuvette inside a Jasco V-770 (Jasco Co., Tokyo, Japan) Peltier

temperature-controlled UV/Vis-NIR scanning spectrophotometer immediately after preparation.

Experiments were performed recording the extinction values at 790 nm every 5 min in a time interval from 0 to 600 min at 25 °C, respectively. Results were plotted against time and fitted using the OriginPro 2021 software with the equation

$$y = y_0 + A_1 e^{\frac{-x}{t_1}} + A_2 e^{\frac{-x}{t_2}} \quad (1)$$

for a two-phase exponential decay function. The half-life parameters were calculated from the following formulae

$$(t_{1/2})_1 = t_1 \cdot \ln 2 \text{ and } (t_{1/2})_2 = t_2 \cdot \ln 2. \quad (2)$$

Adsorption experiments

25 mg of powdered materials (GO or GO-Poly- β CD) were sonicated in 25 mL of tap water for 2 hours. A mixture containing ten PFAS (125 μ L, 100 μ g/L) was then added to the suspensions to achieve a final concentration of 0.5 μ g/L each in a total volume of 25 mL. Similarly, 25 mg of GAC was directly added to 25 mL of tap water, followed by the addition of 125 μ L of the ten-PFAS mixture (100 μ g/L), resulting in the same final PFAS concentration (0.5 μ g/L). All samples were subjected to gentle agitation for 15 minutes, then centrifuged (10.000 rpm, 10 minutes) and analyzed by UPLC-MS/MS.

PFAS quantification

Samples containing PFAS were analyzed by UPLC-MS/MS (ACQUITY UPLC H-Class PLUS coupled to a XEVO TQS Micro mass detector, Waters). A 1 mL aliquot of each sample was used for automated injection. Chromatographic separation was performed on a Waters Acquity UPLC CSH Phenyl-Hexyl column (1.7 μ m, 2.1 $\times$ 100 mm) with a Waters Isolator Column (2.1 $\times$ 50 mm). The column temperature was maintained at 34 °C, the flow rate was 0.3 mL min⁻¹, and the injection volume was 40 μ L. Total run times were 8 min for PFBA and 21 min for the mixture of eight PFAS. The mobile phase consisted of a biphasic gradient: 2 mM NH₄OAc in ultrapure water:methanol (95:5) as phase A, and 2 mM NH₄OAc in methanol as phase B (for further details, see ESI Section 4).

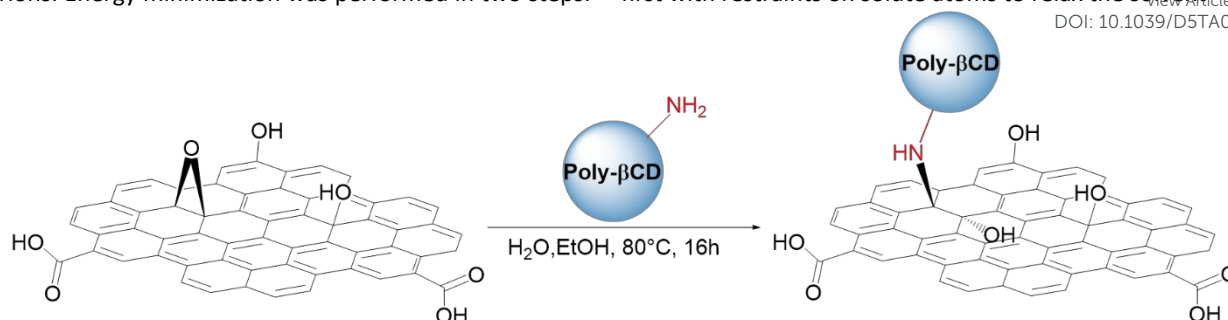
Computational modelling

All MD simulations were carried out using the Amber22 software suite. The PFAS molecules were parameterized using the generalized AMBER force field (GAFF), and atomic partial charges were derived using the restrained electrostatic potential (RESP) fitting procedure after geometry optimization at the HF/6-31G* level of theory. The GO-Poly- β CD model-system was generated by covalently grafting a representative Poly- β CD onto a GO surface, modelled according to experimental XPS data. Different protonation states of the GO-Poly- β CD system were used to mimic pH 4, 7, and 10, where surface groups were protonated or deprotonated accordingly. Complexes of GO-Poly- β CD with PFAS were solvated in an explicit TIP3P water box, ensuring a 20 Å buffer from solute to box edge. The systems were neutralized by adding appropriate



counterions. Energy minimization was performed in two steps: first with restraints on solute atoms to relax the solvent

View Article Online
DOI: 10.1039/D5TA08410E



Scheme 1. Synthetic route to GO-Poly-βCD.

environment, and second without restraints to minimize the entire system. Subsequently, each system was equilibrated under NPT conditions at 298 K for 10 ns using periodic boundary conditions and an Andersen thermostat. Production MD simulations were then run for 100 ns under constant temperature and pressure conditions. To evaluate binding affinities, Molecular Mechanics-Generalized Born Surface Area (MM-GBSA) calculations were performed and the total binding free energy ($\Delta G_{\text{binding}}$) was decomposed into van der Waals (VDW), electrostatic (E_{el}), non-polar solvation (E_{surf}), and entropic ($T\Delta S$) contributions. Entropic contributions were estimated via normal mode analysis. The cavities reported in this study were identified using ChimeraX software (version 1.10.1) with the Cavity tool.

Results and discussion

Synthesis and characterization

The synthesis of GO-Poly-βCD was carried out via epoxide ring-opening reaction of exfoliated GO nanosheets,²⁵ using Poly-βCD-NH₂ obtained by crosslinking of mono-6-azido β-CD in the presence of epichlorohydrin and following reduction of azido in amino groups³¹ (Scheme 1). Under the reaction conditions (80 °C, 16 h), GO nanosheets are partially reduced, thus to clarify the role of Poly-βCD, a control material (GO-Control) was prepared by treating GO under the same conditions as those used for GO-Poly-βCD synthesis, but in the absence of Poly-βCD. The resulting materials were purified from unreacted materials through repeated washing and centrifugation steps with water until the supernatant reached neutral pH. The lyophilized products were then characterized and used for the targeted adsorption experiments. The surface chemical properties of pristine GO, GO-Control, GO-Poly-βCD, and Poly-βCD-NH₂ were investigated by X-ray photoelectron spectroscopy (XPS). Long-range spectra (Figure S1) revealed the presence of carbon, oxygen, and nitrogen in almost all samples; therefore, the corresponding core levels (C1s, N1s, and O1s) were analyzed in detail. Figure 2a-b reports the C1s and N1s core levels of all samples while O1s core level is shown in Figure S2. GO and GO-Control surfaces were found to be almost identical, exhibiting the typical C=C sp² component (285.0 eV), ascribed to the aromatic graphene framework, together with C–O, C=O, and COOR features (287.0–287.1 eV, 288.4–288.7 eV, and 289.3 eV, respectively). The N1s intensity was very low (<1% atomic

concentration) and was detected only in GO-Control. Its deconvolution yielded a C–NH–C signal (graphitic nitrogen) at 400.1 eV, as previously reported,³² and two additional components at 398.5 eV and ~402.0 eV, which can be assigned to pyridinic nitrogen³² and to amino groups²⁵, respectively, although this latter attribution is not straightforward since different N-based functionalities contribute in this range. Poly-βCD-NH₂ displayed a completely different C1s line shape, dominated by C–O bonds at 286.4 eV, and a single N1s peak corresponding to primary amines at 399.8 eV. Upon functionalization of GO, a complete change in the C1s line shape was observed, along with an increase in the intensity of high-binding-energy N1s components. The relative contribution of graphitic and amino-related components in N1s increased from 43.1% in pristine GO to 74.8% in GO-Poly-βCD. These changes are compatible with the covalent functionalization of GO with poly-βCD.^{19, 25} Thermogravimetric analysis (TGA) supported the successful covalent bonding of GO with Poly-βCD. The TGA profile of GO-Poly-βCD (Figure 2c) exhibited a distinct inflection point at 333 °C, clearly visible as a peak in the derivative thermogravimetric (DTG) curve. This thermal event was absent in pristine GO and closely matched the degradation onset of pure Poly-βCD (Figure 2e), which showed a similar DTG peak at 342 °C. The slight shift in decomposition temperature suggested covalent interaction between GO and Poly-βCD, which altered the thermal stability of the polymer moieties. Additionally, the increased weight loss observed in the 200–400 °C range for GO-Poly-βCD compared to pristine GO (Figure 2d) supported the presence of organic functional groups from the grafted polymer. Elemental Analysis (EA) was used to estimate the bulk composition of the synthesised materials and to evaluate the Poly-βCD loading (Table S1). GO-Poly-βCD exhibits a higher nitrogen content compared to pristine GO (0.3% vs 0.08%), ascribed to the presence of amine functionalities in the glucose units. Under the reaction conditions (16 hours at 80 °C), the GO nanosheets tend to undergo partial reduction¹⁵ (GO-Control), leading to a decreased oxygen content (24% vs 30.8%) and, consequently, an increased C/O ratio (1.4 vs 1.3). By monitoring the increase in nitrogen content, it was possible to estimate the Poly-βCD loading. Starting from a βCD/epichlorohydrin ratio of 7:3 in the starting material, the molecular formula of Poly-βCD repetitive unit was determined to be C₆₇H₁₂₉NO₄₂. Given a total of 239 atoms per repetitive unit, the theoretical nitrogen



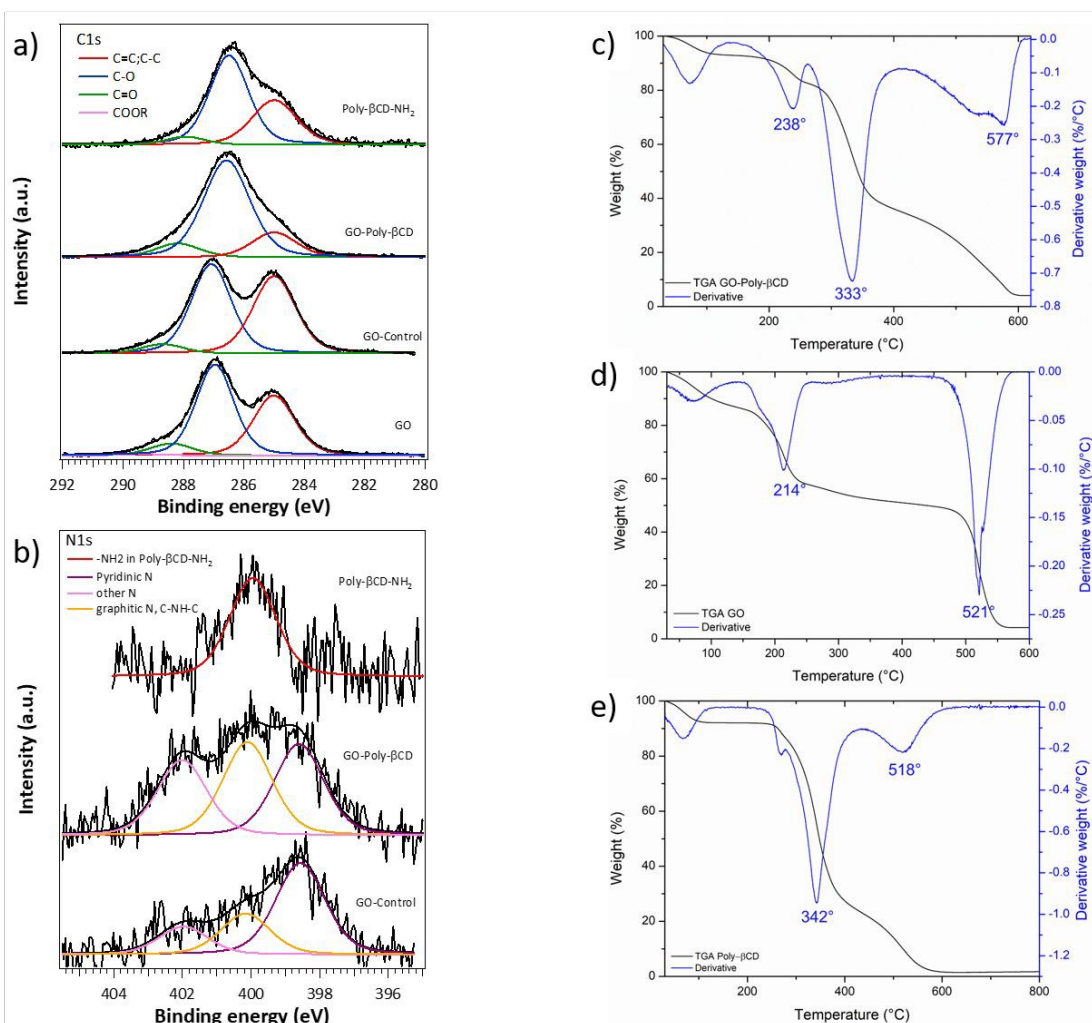


Figure 2. XPS core level spectra of a) C1s and b) N1s of GO, GO Control, GO-Poly-βCD and Poly-βCD-NH₂ (for reference) and TGA profiles of c) GO-Poly-βCD, d) pristine GO and e) Poly-βCD

content was calculated as 0.42%, which is in excellent agreement with the experimentally determined value (Table S1). The Poly-βCD loading in GO-Poly-βCD resulted to be 55% as it can be estimated using Eq. 3, where n_N is the number of Nitrogen atoms involved (one in this case), n_{Total} is the total number of atoms per polymer unit (239) and the % refers to that of atomic nitrogen in GO-Poly-βCD and GO-Control (See supporting info section 1 for more details):

$$L\% = \left[\frac{\%N(GO-Poly-\beta CD) - \%N(GO-Control)}{n_N} \right] \times n_{Total} \quad (3)$$

Figure 3 shows the attenuated total reflection Fourier-transform infrared spectroscopy (ATR-FTIR) spectra of GO-Poly-βCD in comparison with pristine GO and Poly-βCD. All GO-based samples exhibited the broad band in the 3700–3000 cm⁻¹ region, associated with O–H and N–H stretching vibrations (case of GO-Poly-βCD), and the C–H stretching bands of Poly-βCD between 2920 and 2800 cm⁻¹ (case of Poly-βCD and GO-Poly-βCD) confirming the presence of the polymer. In the 1800–1500 cm⁻¹ region, GO-Poly-βCD displayed the characteristic carbonyl

stretching bands of GO at 1710 and 1610 cm⁻¹, indicative of C=O vibrations in carboxylic groups. The band of Poly-βCD at 1636 cm⁻¹ was no longer visible after functionalization, consistent with the chemical modification of the amine groups. Overall, the spectral features confirm the coexistence of both GO and Poly-βCD in the hybrid material.

The colloidal stability of GO-Poly-βCD was assessed by turbidimetry and Dynamic Light Scattering (DLS) (Figure S4–S5, Table S4). Turbidimetric analysis of 250 μg/mL dispersions revealed a two-phase exponential decay with half-lives of ~35 and 883 min, indicating polydispersity with at least two particle populations. DLS at 62.5 μg/mL confirmed a predominant fraction with hydrodynamic diameter, $D_H \approx 100$ nm alongside larger aggregates (~570 nm), with ζ-potential ~–28 mV ensuring stability. pH-dependent studies showed enhanced stability in alkaline conditions, while at pH 2–3 protonation of GO surface groups led to aggregation and increased D_H , evidencing partial destabilization.



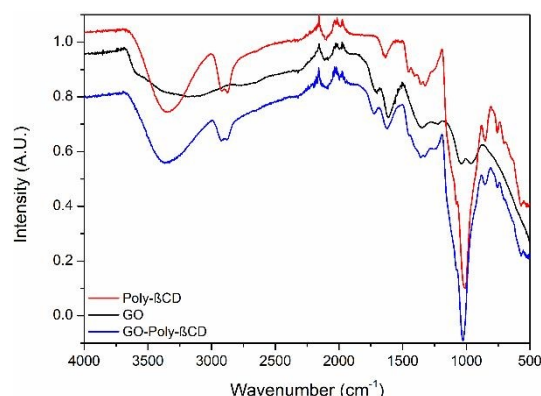


Figure 3. ATR-FTIR spectra of GO (black), Poly-βCD (red) and GO-Poly-βCD (blue)

PFAS adsorption

GO-Poly-βCD nanosheets were tested as sorbents under batch conditions for a mixture of ten PFAS molecules (Figure 4a), varying in chain length ((CF)₃₋₉) and end-group substitution (sulfonate and carboxylate).

The concentration of the PFAS mixture was selected to reflect environmentally relevant levels commonly found in contaminated water sources (0.1–3 μg/L) and the pH of the tested solution was that of tap water (pH 7). To simulate real-field applications, the adsorption contact time was set to 15 minutes, in line with the typical contact time between water and GAC in municipal water treatment plants (approximately 10–20 minutes). The removal of each contaminant was estimated by UPLC-MS-MS analysis (details in ESI, section 4). Figure 4b compares the PFAS removal efficiency of pristine GO, GO-Poly-βCD, and GO-Control after 15 minutes of contact. It can be seen that GO-Poly-βCD outperforms pristine for short- and medium-chain PFAS, achieving up to 97% removal for long-

chain PFAS with (CF)₆₋₉ length. Medium-chain PFAS ((CF)₄₋₆) are also effectively removed by GO-Poly-βCD, with removal efficiencies exceeding 80%. On the other hand, PFPeA ((CF)₄) was adsorbed by GO-Poly-βCD with removal up to 44%, compared to 0% and 11% observed for GO and GO-Control respectively, while PFBA ((CF)₃), one of the most persistent PFAS, was not removed at all. Pristine GO showed negligible adsorption for short- and medium-chain PFAS ((CF)₃₋₇), with increasing removal observed for long-chain PFAS, ranging from 14% to 76%. GO-Control, which exhibits a slightly higher degree of reduction than pristine GO (C/O = 1.4 vs 1.3), showed modest removal of medium-chain PFAS ((CF)₄₋₇) and increasing removal for long-chain PFAS ((CF)₈₋₉), reaching up to 59% for PFDA. Overall, GO-Poly-βCD demonstrates consistently higher adsorption selectivity toward medium-chain PFAS compared to both unmodified GO and GO-Control. This is clearly evidenced by the adsorption of (CF)₄₋₆ chains, which are removed by GO-Poly-βCD with efficiencies ranging from 44% to 97%, whereas GO-Control shows only limited removal, with a maximum of 20% for PFHpA ((CF)₆), and pristine GO exhibits negligible removal across this range. Poly-βCD polymers carry positive charges due to multiple protonated amine moieties, which interact with the negatively charged, deprotonated carboxylate groups of GO. As a result, pH variations can strongly influence the adsorption behavior of the GO-Poly-βCD system. We therefore compared the removal performances at pH 4, 7, and 10 (Figure 4c). Adsorption at pH 4 is consistently higher than at pH 7, with the most pronounced improvements observed for short- and medium-chain PFAS. At pH 4, PFBS, PFHxA, and PFHpA were removed with exceptional efficiencies of 95%, 94%, and 97%, respectively. PFPeA, which exhibited 44% removal at pH 7, showed an increased removal of 80% at pH 4.

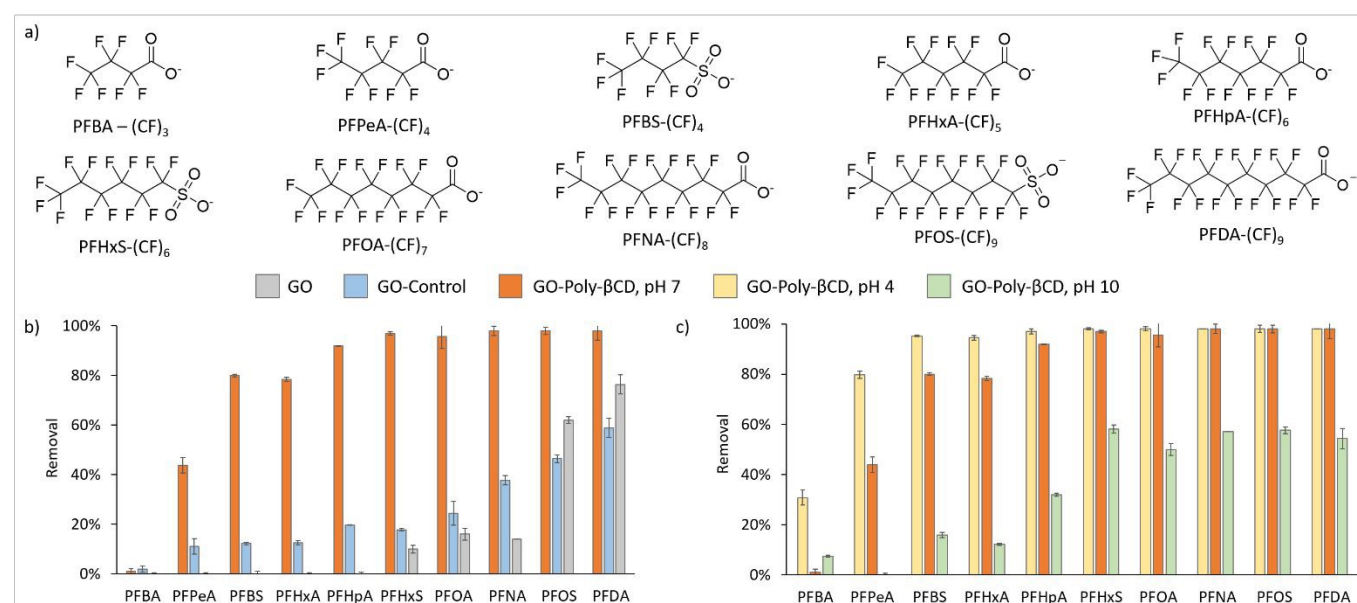


Figure 4. a) Molecular structure of the PFAS molecules, with different chain length ((CF)₃₋₉) and end-group substitution (sulfonate and carboxylate) studied in this work. b) Removal of PFAS mixture at pH 7 (orange) from GO-Poly-βCD compared to GO-Control (blue) and unmodified GO (grey). c) Experiments at pH 4 (yellow), pH 7 (orange) and pH 10 (green).



Notably, PFBA, which was not removed at neutral pH, was adsorbed at acidic pH with a removal efficiency of 31%. Under basic conditions, the system exhibited substantially reduced adsorption: short- and medium-chain PFAS ((CF)₃₋₆) were removed with negligible to poor efficiency, while long-chain PFAS ((CF)₇₋₉) showed only moderate adsorption, with a maximum removal of 58% for PFHxS, significantly lower than the near-quantitative removal observed under neutral and acidic conditions.

Adsorption mechanistic insights

To elucidate the binding mechanism responsible for the enhanced removal of PFAS by GO-Poly- β CD at pH 7, molecular dynamics (MD) simulations were performed using perfluorohexanoic acid (PFHxA) as a representative case study. To investigate the specific role of Poly- β CD moieties grafted onto the GO surface, the interactions of PFHxA with both GO and GO-Poly- β CD were analyzed. Five distinct adsorption sites for PFHxA were identified in the GO-Poly- β CD system (Figure 5), designed to capture a range of interaction environments: adsorption on GO (Case 1), interaction with Poly- β CD (Case 2), encapsulation within the cyclodextrin cavity (Case 3), and sandwich interaction with both GO and Poly- β CD (Case 4). Five independent MD simulations, each 100 ns in length, were carried out from these starting geometries to dynamically sample the interactions between PFHxA and the nanosorbent. The binding free energy between PFHxA and GO-Poly- β CD was calculated in each case using the MM-GBSA approach (Table 1). PFHxA exhibited weak binding to GO ($\Delta G_{\text{binding}} = -1.9 \text{ kcal mol}^{-1}$ for pristine GO, $-2.8 \text{ kcal mol}^{-1}$ for GO-Poly- β CD, Case 1). Similarly, interactions between PFHxA and Poly- β CD alone were very weak (Case 2, $-1.8 \text{ kcal mol}^{-1}$).

Encapsulation of PFHxA within the β CD cavity of the Poly- β CD polymer (Case 3) significantly enhanced binding ($\Delta G_{\text{binding}} = -12.4 \text{ kcal mol}^{-1}$). Notably, the strongest interaction was observed in Case 4, where PFHxA is sandwiched between Poly- β CD and GO, interacting with both components of the nanosorbent, yielding a $\Delta G_{\text{binding}}$ of $-15.6 \text{ kcal mol}^{-1}$. This value demonstrates that Poly- β CD is more

effective than monomeric β CD derivatives (PFBA: β CD = $-12.3 \text{ kcal mol}^{-1}$).²⁵ The high affinity arises from a favorable combination of interaction forces, primarily driven by van der Waals interactions, which contribute $-35.2 \text{ kcal mol}^{-1}$. These strong dispersion forces result from the close contact between PFHxA and both the GO surface and the overlying Poly- β CD. The electrostatic contribution is slightly unfavorable ($E_{\text{el}} = 8.0 \text{ kcal mol}^{-1}$), likely due to repulsion between the negatively charged head group of PFHxA and the GO surface, as well as desolvation of the PFAS head group.

In contrast, non-polar solvation contributions were favorable ($E_{\text{SURF}} = -5.1 \text{ kcal mol}^{-1}$), reflecting burial of the PFAS aliphatic chain within the cleft formed by Poly- β CD and the GO sheet. The entropy term ($\Delta S = -16.7 \text{ kcal mol}^{-1}$) represents a binding penalty due to the conformational restriction of PFHxA upon adsorption. These results indicate that stable adsorption and retention of PFAS require the combined action of both Poly- β CD and GO moieties, highlighting the necessity of both components for effective pollutant capture and supporting a cooperative “sandwich-like” mechanism as the most favorable binding mode. The binding energies of all investigated PFAS on GO-Poly- β CD were also calculated, reproducing the experimentally observed trends (Figure S3).

To assess the influence of pH on PFAS adsorption, MD simulations were performed for a model system (PFHxA adsorption on GO-Poly- β CD) at pH 4, 7, and 10, and the corresponding binding free energies were calculated (Table 2). A net decrease in binding affinity was observed with increasing pH: $\Delta G_{\text{binding}} = -19.9$, -15.6 , and $-10.6 \text{ kcal mol}^{-1}$ at pH 4, 7, and 10, respectively, consistent with the experimental removal trend (Figure 4c). Electrostatic repulsion (E_{el}) between PFHxA and GO-Poly- β CD increases as pH rises, since PFHxA remains deprotonated across the studied range ($\text{pK}_a = -0.16$). The improved adsorption under acidic conditions is primarily attributed to stronger van der Waals interactions, which are unexpectedly influenced by pH. Structural analysis revealed a pH-dependent conformational change of the nanosystem, affecting the accessibility and stability of PFAS adsorption sites (Figure 6). At pH 4, Poly- β CD remains closely associated with GO

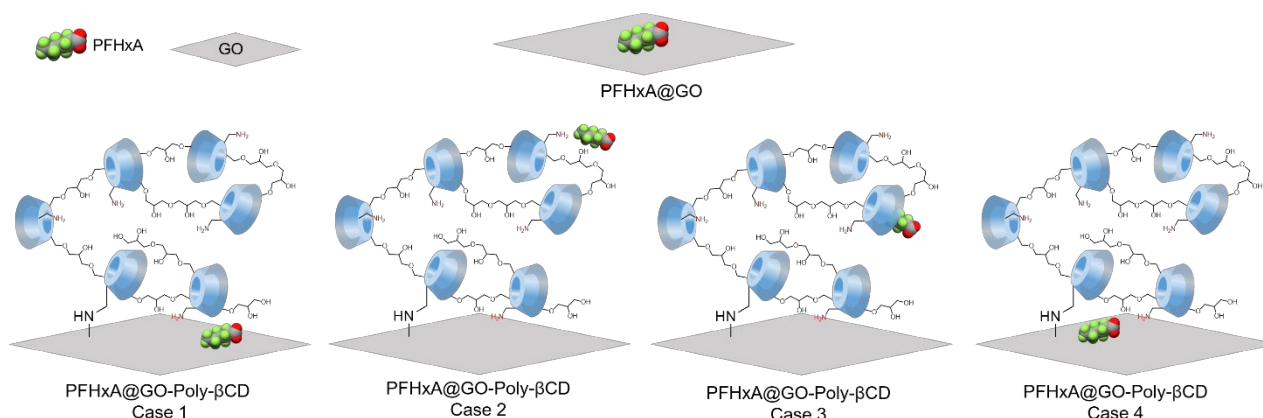


Figure 5. Schematic representation of the interactions between PFHxA and GO-Poly- β CD, highlighting the different adsorption sites.



Table 1. Binding free energy ($\Delta G_{\text{binding}}$) and energy contributes calculated for each material. All energies are reported in kcal mol⁻¹.

PFHxA@	VDW	E _{El}	E _{SURF}	$\Delta H_{\text{binding}}$	ΔS	$\Delta G_{\text{binding}}$
GO	-16.0	4.8	-1.2	-12.4	-10.5	-1.9
Case 1	-18.7	4.2	-1.4	-15.9	-13.1	-2.8
Case 2	-15.2	4.4	-2.7	-13.4	-11.6	-1.8
Case 3	-29.7	6.0	-4.3	-28.0	-15.6	-12.4
Case 4	-35.2	8.0	-5.1	-32.3	-16.7	-15.6

through hydrogen bonding, forming tight binding clefts (365.04 Å²) for PFHxA.

These clefts become less well-defined at pH 7 (443.94 Å²), while at pH 10, Poly-βCD partially detaches from the GO surface, disrupting the interaction network and dismantling the clefts responsible for PFHxA adsorption. Consequently, under basic conditions, PFHxA remains largely exposed on the GO surface and is only partially covered by the polymer, breaking the sandwich-like adsorption site, reducing the synergistic binding effect, and resulting in overall weaker interactions. This indicates that the adsorption of PFHxA—and PFAS in general—on GO-Poly-βCD is primarily hydrophobically driven. This behavior is also reflected in the non-polar solvation term (E_{SURF}), which becomes less stabilizing at higher pH (E_{SURF} = -3.2 kcal mol⁻¹ at pH 10 vs. -5.1 kcal mol⁻¹ at pH 4). As with van der Waals interactions, E_{SURF} depends directly on the contact area between the adsorption site and the PFAS, i.e., their shape complementarity.

Although it was initially hypothesized that pH would primarily modulate electrostatic interactions between PFAS and the GO-

Poly-βCD surface—particularly involving the protonated amines of Poly-βCD—our findings indicate that the critical factor is the pH-dependent conformational rearrangement of the hybrid material. At lower pH, stronger interactions between GO and Poly-βCD generate well-defined clefts that facilitate efficient guest inclusion and encapsulation of PFAS.

To further confirm the pivotal role of Poly-βCD in modulating the adsorption behavior of the hybrid material under different pH conditions, the removal performance of pristine GO was also evaluated across pH 4, 7, and 10 (Figure S6). Short- and medium-chain PFAS (CF₃₋₆) exhibited negligible removal at all tested pH values. A modest improvement was observed at pH 4, with removal increasing to 16% for PFHpA and 29% for PFHxS. This can be attributed to partial neutralization of the negative surface charge of GO under acidic conditions, slightly reducing electrostatic repulsion with the anionic PFAS. For long-chain PFAS (CF₇₋₉), removal efficiency increased more substantially under acidic conditions, reaching up to 91% in the case of PFDA. At neutral and basic pH, only minor improvements were observed for PFNA, while PFOS and PFDA showed higher removal efficiencies, though never exceeding 80%. The gradual increase in adsorption with chain length across all pH conditions reflects the stronger hydrophobic interactions of longer-chain PFAS. Overall, the negligible adsorption of pristine GO under most conditions highlights the essential role of Poly-βCD in stabilizing PFAS and corroborates the experimental and computational findings described above.

PFAS Adsorption of GO-Poly-βCD vs GAC

To evaluate the potential of GO-Poly-βCD for drinking water treatment, its adsorption performance was compared with that of GAC, representing the current industrial sorbent benchmark

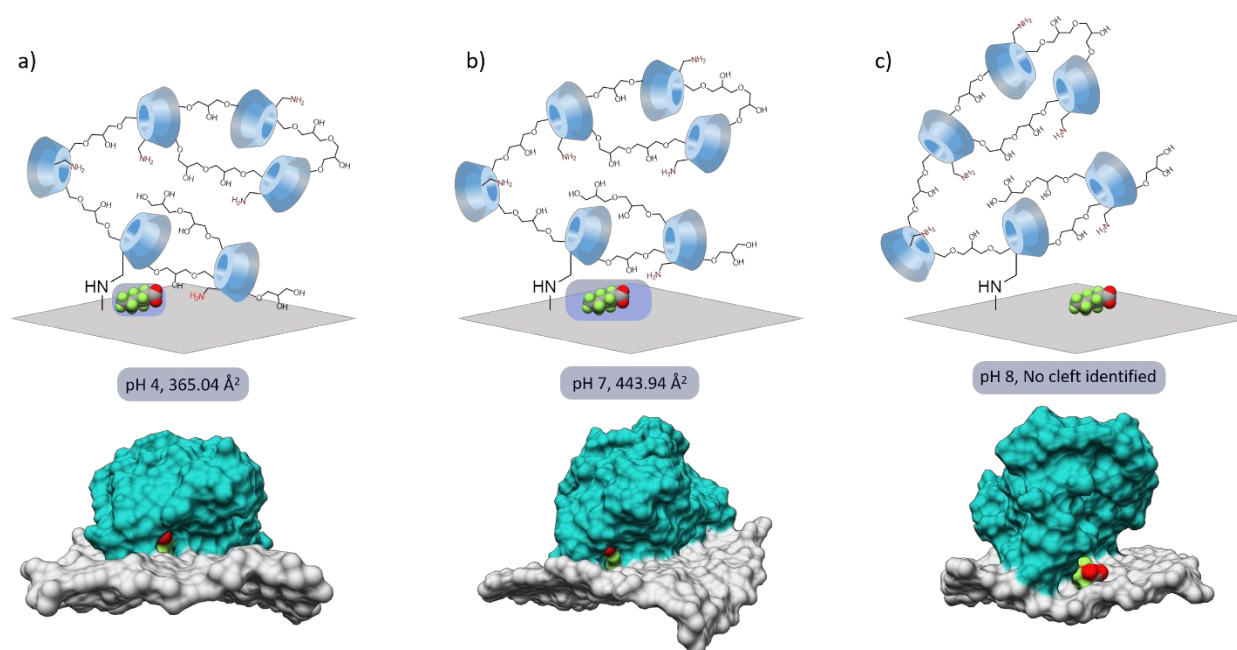
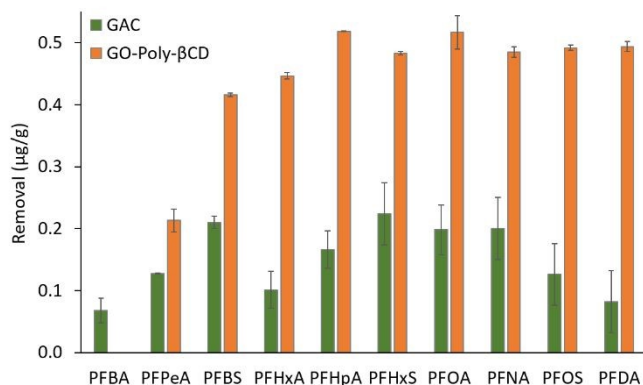
**Figure 6.** Sketch of PFHxA@GO-Poly-βCD and relative cleft formation and calculated area at a) pH 4, b) pH 7 and c) pH 10.

Table 2. Binding free energies ($\Delta G_{\text{binding}}$) obtained for PFHxA@GO-Poly- β CD at three different pH (4, 7 and 10). All energies are reported in kcal mol⁻¹.

PFHxA@GO-Poly- β CD	VDW	E_{EI}	E_{SURF}	$\Delta H_{\text{binding}}$	ΔS	$\Delta G_{\text{binding}}$
pH 4	-37.8	7.8	-5.1	-35.2	-15.3	-19.9
pH 7	-35.2	7.9	-5.1	-32.3	-16.7	-15.6
pH 10	-31.2	10.6	-3.2	-23.7	-13.1	-10.6

**Figure 7.** PFAS removal performances of GO-Poly- β CD and GAC.

widely employed in treatment plants for tertiary and quaternary water treatment as well as in point-of-use purification systems.

The comparison was performed in tap water at neutral pH, with contact time of 15 minutes, conditions that closely reflect real-world applications. Figure 7 shows the removal efficiencies of GO-Poly- β CD and GAC, expressed as micrograms of PFAS removed per gram of sorbent ($\mu\text{g g}^{-1}$). The data clearly indicate that GO-Poly- β CD outperforms GAC in PFAS adsorption. While GAC achieves a maximum removal of $0.2 \mu\text{g g}^{-1}$ for most of the selected PFAS, GO-Poly- β CD reaches up to $0.5 \mu\text{g g}^{-1}$ within the same contact time. For PFPeA ((CF)₄), GAC removes $0.1 \mu\text{g g}^{-1}$, slightly lower than the $0.2 \mu\text{g g}^{-1}$ achieved by GO-Poly- β CD. The shortest-chain PFAS, PFBA, is only minimally adsorbed by GAC ($0.07 \mu\text{g g}^{-1}$) and remains undetectable for GO-Poly- β CD. The advantage of GO-Poly- β CD becomes even more pronounced when considering total PFAS removal, which reaches $4 \mu\text{g g}^{-1}$ compared to $1.3 \mu\text{g g}^{-1}$ for GAC. In addition to higher removal efficiencies, GO-Poly- β CD exhibits rapid adsorption kinetics, achieving significant PFAS uptake within just 15 min. This feature is particularly relevant for practical drinking water purification, including both household point-of-use filters and industrial treatment systems operating under short contact times.

Conclusions

In conclusion, we report the synthesis of a novel GO material covalently grafted with poly- β -cyclodextrin crosslinked with epichlorohydrin. The resulting hybrid is insoluble in water, unlike previously reported β -cyclodextrin-based polymers that often exhibit partial solubility, making it more water-processable and suitable for practical applications in aqueous

environments. GO-Poly- β CD material adsorption performances were evaluated on a mixture of PFAS with varying chain lengths ((CF)₃₋₉) and acidic and sulfonic functional groups. Batch adsorption experiments demonstrated that GO-Poly- β CD outperforms pristine GO in the removal of medium-chain PFAS, with remarkable efficiencies of 97% for PFHxA and PFHpA ((CF)₆₋₇) and 80% for PFPeA ((CF)₄), whereas pristine GO and GO-Control exhibited negligible adsorption in this range. Compared to GAC, GO-Poly- β CD achieved higher overall PFAS uptake within 15 minutes of contact, highlighting its superior performance under field-relevant conditions. Molecular dynamics simulations revealed that PFAS adsorption is governed by a synergistic mechanism, in which contaminants are stabilized through a pH-responsive conformational arrangement of GO grafted with poly- β CD. Overall, these findings provide new insights into the design of hybrid nanomaterials for PFAS remediation and underscore the potential of poly- β -cyclodextrin-based GO materials as efficient and sustainable sorbents for drinking water treatment.

Author contributions

Andrea Trifoglio: methodology, investigation, reviewing editing, writing original draft. Angela Pintus: methodology, investigation. Sara Khaliha: methodology, investigation. Marco Agnes: methodology. Lucrezia Aversa: methodology, investigation. Roberto Verucchi: methodology, investigation. Nina Burduja: investigation, Giuseppe Nocito: methodology, investigation and review original draft, Antonino Mazzaglia: methodology and review original draft. Tainah Dorina Marforio: investigation, methodology, formal analysis. Matteo Calvaresi: methodology, investigation. Manuela Melucci: coordination, conceptualization, validation, writing original draft.

Conflicts of interest

There are no conflicts to declare.

Data availability

Supporting Information. GO-poly β CD loading calculation, X-Ray Photoelectron spectroscopy (XPS), Binding affinities, UPLC-MS/MS details, Colloidal stability studies, and PFAS adsorption from GO at different pH are available free of charge at <https://pubs.rsc.org/>

Acknowledgements

MM and MA thanks Dr. I. Manet and Dr A. Kovtun for their scientific support and Cyclolab for kindly providing Poly- β CD-NH₂. The authors gratefully acknowledge the support of this work by the projects Life-Remembrance, ENV/IT/001001 Life Resource and Environment LIFE20 'Give plastic wastes from the production of hollow-fiber membranes a second life', project PNRR MUR project ECS_00000033_ECOSISTER-ATOS and project Capitale Naturale e Risorse per il futuro dell'Italia (FOE



ARTICLE

Journal Name

2020, DTA.AD005.314). M.A. thanks European Union's Horizon Europe Research and Innovation Programme under the MSCA grant agreement Bicyclos N° 101130235. Views and opinions expressed are however those of the author(s) only and do not necessarily reflect those of the European Union. TDM was supported by Fondazione Umberto Veronesi.

Notes and references

1. *Nature Water*, 2023, **1**, 993-993.
2. O. S. Arvaniti and A. S. Stasinakis, *Science of The Total Environment*, 2015, **524-525**, 81-92.
3. M. Sun, E. Arevalo, M. Strynar, A. Lindstrom, M. Richardson, B. Kearns, A. Pickett, C. Smith and D. R. U. Knappe, *Environmental Science & Technology Letters*, 2016, **3**, 415-419.
4. S. Garg, J. Wang, P. Kumar, V. Mishra, H. Arafat, R. S. Sharma and L. F. Dumée, *Journal of Environmental Chemical Engineering*, 2021, **9**, 105784.
5. J. Gardiner, *Australian Journal of Chemistry*, 2015, **68**, 13-22.
6. S. Dolui, D. Kumar, S. Banerjee and B. Ameduri, *Accounts of Materials Research*, 2021, **2**, 242-251.
7. Z. Du, S. Deng, Y. Bei, Q. Huang, B. Wang, J. Huang and G. Yu, *Journal of Hazardous Materials*, 2014, **274**, 443-454.
8. T. H. Boyer, Y. Fang, A. Ellis, R. Dietz, Y. J. Choi, C. E. Schaefer, C. P. Higgins and T. J. Strathmann, *Water Research*, 2021, **200**, 117244.
9. P. McCleaf, S. Englund, A. Östlund, K. Lindegren, K. Wiberg and L. Ahrens, *Water Research*, 2017, **120**, 77-87.
10. S. Khaliha, A. Bianchi, A. Kovtun, F. Tunioli, A. Boschi, M. Zambianchi, D. Paci, L. Bocchi, S. Valsecchi, S. Polesello, A. Liscio, M. Bergamini, M. Brunetti, M. Luisa Navacchia, V. Palermo and M. Melucci, *Separation and Purification Technology*, 2022, **300**, 121826.
11. E. Gagliano, M. Sgroi, P. P. Falciglia, F. G. A. Vagliasindi and P. Rocco, *Water Research*, 2020, **171**, 115381.
12. A. M. Kennedy, A. M. Reinert, D. R. U. Knappe, I. Ferrer and R. S. Summers, *Water Research*, 2015, **68**, 238-248.
13. S. Khaliha, F. Tunioli, L. Foti, A. Bianchi, A. Kovtun, T. D. Marforio, M. Zambianchi, C. Bettini, E. Briñas, E. Vázquez, L. Bocchi, V. Palermo, M. Calvaresi, M. L. Navacchia and M. Melucci, *Environmental Science: Water Research & Technology*, 2024, **10**, 1097-1107.
14. S. Khaliha, T. D. Marforio, A. Kovtun, S. Mantovani, A. Bianchi, M. Luisa Navacchia, M. Zambianchi, L. Bocchi, N. Boulanger, A. Iakunkov, M. Calvaresi, A. V. Talyzin, V. Palermo and M. Melucci, *FlatChem*, 2021, **29**, 100283.
15. S. Guo, S. Garaj, A. Bianco and C. Ménard-Moyon, *Nature Reviews Physics*, 2022, **4**, 247-262.
16. M. Melucci, L. Bocchi, M. Zambianchi and V. Palermo, *Nature Water*, 2025, **3**, 369-371.
17. S. Mantovani, S. Khaliha, L. Favaretto, C. Bettini, A. Bianchi, A. Kovtun, M. Zambianchi, M. Gazzano, B. Casentini, V. Palermo and M. Melucci, *Chemical Communications*, 2021, **57**, 3765-3768.
18. S. Mantovani, S. Khaliha, T. D. Marforio, A. Kovtun, L. Favaretto, F. Tunioli, A. Bianchi, G. Petrone, A. Liscio, V. Palermo, M. Calvaresi, M. L. Navacchia and M. Melucci, *Chemical Communications*, 2022, **58**, 9766-9769.
19. A. Trifoglio, S. Mantovani, S. Khaliha, A. Kovtun, T. D. Marforio, M. Calvaresi and M. Melucci, *Nanoscale*, 2025, **17**, 12124-12133.
20. G. Crini, *Chemical Reviews*, 2014, **114**, 10940-10975.
21. A. H. Karoyo, A. S. Borisov, L. D. Wilson and P. Hazendonk, *The Journal of Physical Chemistry B*, 2011, **115**, 9511-9527.
22. P. Lo Nostro, I. Santoni, M. Bonini and P. Baglioni, *Langmuir*, 2003, **19**, 2313-2317.
23. H. Tatsuno and S. Ando, *The Journal of Physical Chemistry B*, 2006, **110**, 25751-25760.
24. A. H. Karoyo, P. S. Sidhu, L. D. Wilson, P. Hazendonk and A. Borisov, *The Journal of Physical Chemistry C*, 2015, **119**, 22225-22243.
25. F. Tunioli, T. D. Marforio, L. Favaretto, S. Mantovani, A. Pintus, A. Bianchi, A. Kovtun, M. Agnes, V. Palermo, M. Calvaresi, M. L. Navacchia and M. Melucci, *Chemistry – A European Journal*, 2023, **29**, e202301854.
26. Y. Miyah, N. El Messaoudi, M. Benjelloun, J. Georgin, D. S. P. Franco, M. El-habacha, O. A. Ali and Y. Acikbas, *Carbohydrate Polymers*, 2025, **350**, 122981.
27. L. Xiao, Y. Ling, A. Alsbaiee, C. Li, D. E. Helbling and W. R. Dichtel, *Journal of the American Chemical Society*, 2017, **139**, 7689-7692.



28. A. Alsbaiee, B. J. Smith, L. Xiao, Y. Ling, D. E. Helbling and W. R. Dichtel, *Nature*, 2016, **529**, 190-194.
29. J.-L. Sun, H. Zeng and H.-G. Ni, *Chemosphere*, 2013, **90**, 1751-1759.
30. C. Belenguer-Sapiña, E. Pellicer-Castell, A. R. Mauri-Aucejo, E. F. Simó-Alfonso and P. Amorós, *Journal*, 2021, **11**.
31. A.-M. Hada, M. Potara, S. Astilean, A. Cordaro, G. Neri, M. Malanga, A. Nostro, A. Mazzaglia, A. Scala and A. Piperno, *Carbohydrate Polymers*, 2022, **293**, 119736.
32. X. Deng, A. Verdaguer, T. Herranz, C. Weis, H. Bluhm and M. Salmeron, *Langmuir*, 2008, **24**, 9474-9478.

View Article Online
DOI: 10.1039/D5TA08410E



[View Article Online](#)

DOI: 10.1039/D5TA08410E

Data availability

Electronic Supplementary Information. X-Ray Photoelectron spectroscopy (XPS), binding affinities, UPLC-MS/MS, Colloidal stability, PFAS adsorption from GO at different pH “This material is available free of charge via the Internet at [http:// https://rsc.org](http://https://rsc.org).”

