



Cite this: *React. Chem. Eng.*, 2024, 9, 888

Interaction of SO₂ with a Cu–Mn oxide oxygen carrier during chemical looping with oxygen uncoupling†

Turna Barua^a and Bihter Padak^b ^{*}

Chemical looping with oxygen uncoupling (CLOU) is a variant of the chemical looping combustion (CLC) process, where fuel combustion occurs *via* metal oxides, known as oxygen carriers. The purpose of this study is to investigate how SO₂ interacts with a Cu–Mn oxide oxygen carrier during CH₄ combustion in the CLOU process. The oxygen carrier was reduced *via* CH₄–N₂ gas mixture with and without SO₂ in a batch fluidized-bed reactor. The results indicate that SO₂ does not affect the oxygen release capacity of a Cu–Mn oxide. However, it negatively affects the CH₄ conversion by increasing CO formation as well as the unreacted CH₄ amount. This adverse effect becomes more pronounced as the SO₂ concentration increases. Increasing the temperature from 850 to 950 °C improves the conversion of CH₄. During a reduction cycle, SO₂ oxidizes to SO₃, creating a competing demand for oxygen between fuel and SO₂, and lowers CH₄ conversion if there is not enough oxygen for both reactions. Hence, decreasing the feed CH₄ concentration with a fixed oxygen carrier amount improves CH₄ conversion in the presence of SO₂. Additionally, characterization of SO₂-exposed oxygen carriers shows that sulfate species form on the surface, possibly causing a further reduction in CH₄ conversion. However, the presence of SO₂ does not cause a permanent deactivation of the particles and they can be fully regenerated during the oxidation cycle.

Received 22nd September 2023,
Accepted 21st December 2023

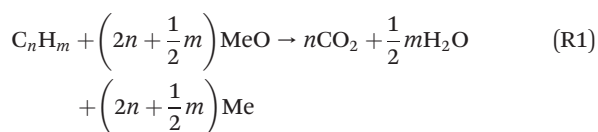
DOI: 10.1039/d3re00498h

rsc.li/reaction-engineering

Introduction

Fossil-fuel combustion accounts for about 60% of total electricity in the U.S. despite increased renewable energy resources.¹ Burning fossil fuels generates a considerable amount of CO₂, a significant greenhouse gas. Carbon capture and storage (CCS) is a promising technology to limit CO₂ concentration in the atmosphere. Since a high purity CO₂ stream is essential to ensure the economic viability of CCS technologies, several carbon-capture methods have been proposed and investigated in recent years.^{2–5} Of these methods, chemical looping combustion (CLC) is a promising technology because of its inherent ability to produce N₂-free CO₂ during fuel combustion while bypassing an expensive gas separation method. In CLC, fuel oxidation occurs *via* a metal oxide, also referred to as an oxygen carrier, instead of air, *via* reaction (R1), therefore avoiding direct contact between fuel and air. As a result, after steam condensation, a highly concentrated CO₂ stream is obtained. The reduced oxygen

carrier is regenerated by air *via* (R2) before being transferred back to the fuel reactor.^{6–12}



The oxidation by which a metal oxide's lattice oxygen reacts with the fuel is the primary distinguishable characteristic of chemical looping technologies. This phenomenon may occur in two ways: i) surface reactions between the fuel and lattice oxygen, (R1), and ii) reaction with gaseous oxygen, (R4), from metal oxide decomposition, (R3), at suitable temperature and oxygen partial pressures. The latter mechanism, where fuel reacts with gaseous oxygen like the traditional combustion reaction, is commonly referred to as chemical looping with oxygen uncoupling (CLOU). The CLOU mechanism is

^a Department of Chemical and Biomolecular Engineering, University of California, Irvine, USA

^b Department of Mechanical and Aerospace Engineering, University of California, Irvine, USA. E-mail: bpadak@uci.edu

† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3re00498h>

particularly beneficial for solid fuel utilization due to a faster gas–solid reaction rate relative to a solid–solid reaction in CLC.^{7,13–22}

Early CLOU research identified monometallic CuO/Cu₂O, Mn₂O₃/Mn₃O₄ oxide-systems as suitable CLOU oxygen carriers. However, these oxides have some limitations like sintering of copper oxides and slow oxidation kinetics of Mn₃O₄ to Mn₂O₃.²³ Therefore, to overcome this limitation, copper and manganese oxides are combined with inert oxides such as TiO₂, ZrO₂, SiO₂, or active metal oxides such as NiO, CaO, and Fe₂O₃.^{14,24–27} One such oxygen carrier is bimetallic Cu–Mn oxide, which utilizes the oxygen uncoupling characteristics of CuO and Mn₂O₃ and negates the disadvantages of constituent oxides.²⁸ It transitions between spinel Cu_xMn_{3–x}O₄ and CuMnO₂ during oxygen uncoupling that provides a considerable amount of gaseous O₂ that results in high reactivity with the fuel. Furthermore, CuMnO₂ can also react with the fuel *via* heterogeneous reactions (CLC) to form Cu and MnO.^{29–31} Previous studies employing a Cu–Mn oxide achieved high combustion efficiency for both solid and gaseous fuels.^{28,32–37} However, the investigations mainly involved sulfur-free or low-sulfur fuels to highlight the redox reactivity of the oxygen carrier. Although one study tested high-sulfur coal (5.2 wt% S) with a Cu–Mn oxygen carrier, it did not investigate the potential impact that sulfur contamination may have on the carrier.³⁸ Given the variable presence of sulfur in fuels such as coal (200–10 000 ppm H₂S in coal-derived synthesis gas) and natural gas (up to 20 ppm H₂S), the oxygen carrier will most likely be exposed to sulfur-containing species in the fuel reactor.^{39,40}

Sulfur is found in coal in various forms, including sulfides, sulfates, and organic sulfur compounds and most part of the coal-S is released during combustion from sulfides and organic bound sulfur.⁴¹ At temperatures higher than 1000 °C and oxygen-rich conditions, SO₂ is the thermodynamically favored species. H₂S can be formed in the gas phase or released from coal-S under substoichiometric conditions. At lower temperatures, the equilibrium shifts towards SO₃ and typically 0.1 to 1% of the SO₂ is oxidized to SO₃ during air combustion *via* reactions (R5)–(R7) in the gas phase.⁴²



SO₃ formation is influenced by the temperature profile, residence time, concentrations of SO₂ and O₂, fly ash composition and presence of catalysts *e.g.* Fe₂O₃.⁴³ In the CLOU process, with the presence of the oxygen carrier acting as an oxidizing catalyst, SO₂ could be oxidized to SO₃ heterogeneously. Due to the likelihood of the interactions of sulfur species and the oxygen carrier, understanding how sulfur species affect the oxygen carrier's performance is

crucial to developing the chemical looping technology. Nevertheless, studies looking into the impacts of sulfur on Cu and Mn-based oxygen carriers during the CLOU process are limited. A study investigating high-sulfur lignite combustion in a continuous CLOU reactor with CuO–MgAl₂O₄ observed SO₂ as the major sulfur species in the effluent. In this study, the sulfur mass balance closure was ~80 wt%, with the rest remaining unaccounted. The authors assumed that the remaining sulfur was accumulated on the oxygen carrier; however, they did not observe a decrease in the reactivity of the oxygen carrier due to the presence of sulfur.⁴⁴ In a similar CLOU study with a Cu-based oxygen carrier (CuO–Fe₂O₃–MgAl₂O₄), SO₂ was also observed as the final sulfur product.⁴⁵ They observed no decrease in the reactivity of the oxygen carrier when exposed to sulfurous fuel; however, the oxygen uncoupling capacity decreased ~10% after 35 hours of operation. Around 15 wt% of the total sulfur was unaccounted for in the sulfur mass balance. Through a controlled TGA study followed by SEM-EDX analysis, (CuO)·(CuSO₄) formation was observed at 930 °C even though it is thermodynamically unstable above 700 °C. The authors speculated that combination of different oxides in the oxygen carrier causes a synergic effect making its reaction affinity with SO₂ different from those of the individual oxides.⁴⁵ Besides Cu-based oxides, a study with a Mn–Si oxide investigated sulfur effects by introducing 5000 ppm SO₂ with CH₄.⁴⁶ The oxygen carrier displayed no appreciable decrease in reactivity or oxygen uncoupling capacity at a maximum exposure of 5000 ppm SO₂.

On the other hand, Cu and Mn-based oxides tend to form sulfides in the CLC process depending on the operating conditions.^{39,47–51} For Cu-based oxygen carriers, thermodynamic calculations show that the possibility of sulfide formation is high at oxygen-deficient conditions, while it is less likely to form at oxygen-rich conditions. For a Mn-based oxide, MnSO₄ formation is thermodynamically possible at both oxygen-deficient and rich conditions.⁴⁹ A CLOU fuel reactor usually operates in an oxygen-rich environment which decreases the possibility of sulfide formation.⁵² However, local regions with high reducing potentials may promote sulfide formation.

Given the prospects of Cu–Mn oxide as an oxygen carrier in CLOU and the potential probability of sulfur interaction with constituent elements, it is necessary to examine the influence of sulfur on this oxygen carrier. Therefore, the objective of this study is to evaluate how sulfur affects the Cu–Mn oxygen carrier in a CLOU process. The study simulates coal combustion in a CLOU process by introducing CH₄ as fuel since it is released during the volatilization stage of coal combustion. Since SO₂ is the major sulfur species formed during solid fuel combustion in CLOU,⁴⁵ it was added as a sulfur impurity. Effects of SO₂ on the oxygen carrier's reactivity with CH₄ was investigated in a laboratory-scale fluidized bed reactor. Finally, Cu–Mn oxygen carrier particles were characterized before and after SO₂ exposure to determine its interaction with SO₂.

Experimental

Materials

The bi-metallic Cu–Mn oxide, synthesized by the incipient wetness impregnation method,⁵³ consists of 34 wt% CuO and 66 wt% Mn₃O₄. Before impregnation, Mn₃O₄ was mixed with graphite (5 wt% of Mn₃O₄) and sintered at 950 °C for one hour. A certain amount of copper nitrate solution (60–80 °C) corresponding to the pore volume of Mn₃O₄ was added to the sintered particles. The impregnated samples were dried at 70 °C overnight and were calcined at 600 °C in an air atmosphere for two hours. The desired loading was achieved by successive 3–4 impregnation steps followed by drying and calcination. Finally, the particles were heated at 950 °C for 6 hours at 10 °C min^{−1} and sieved to a size of 150–300 μm. The final oxygen carrier obtained is referred to as Cu34Mn66 in this study.

Characterization of the oxygen carrier

The powder X-ray diffraction (XRD) method (Rigaku Ultima-III; 30 kV and 40 mA) with Cu-Kα radiation was used to determine the crystalline structure of the oxygen carrier. The pattern was collected in the range of 10 to 90°. The X-ray photoelectron spectroscopy (XPS) analysis was conducted with an AXIS Supra by Kratos Analytical (300 W, 20 mA) equipped with a hemispherical electron analyzer and monochromatic Al-Kα source to analyze the surface chemical composition. Survey and high-resolution spectra of elements were acquired with 160 and 40 eV pass energies, respectively. The binding energies were calculated referencing the C 1s peak energy of 284.8 eV. Temperature programmed reduction (TPR) *via* hydrogen was performed using a Micromeritics 3Flex Surface and Catalyst Analyzer. A 10% H₂/Ar mixture of 30 ml min^{−1} flow rate was used to reduce 60 mg particles. Before reduction, particles were pretreated at 200 °C in N₂ for 1 hour, then cooled to 50 °C. The temperature was ramped up to 800 °C at a heating rate of 10 °C min^{−1}. Isopropanol and liquid N₂ slurry were used to condense the water formed during the process. A thermal conductivity detector (TCD) was used to measure the H₂ consumption signal. Thermogravimetric analysis (TGA) was utilized to determine the oxygen release capacity of Cu34Mn66 particles. Netzsch TG 209 F1 Libra was used to analyze the oxygen release capacity at the following isothermal conditions: 850 °C, 900 °C, and 950 °C. Around 7–9 mg sample was reduced in N₂ and oxidized in air at a 180 ml min^{−1} flow rate. The oxygen uncoupling capacity (R_o) was determined using eqn (E1):

$$R_o = (m_{ox} - m_{red})/m_{ox} \quad (E1)$$

Here m_{ox} is the weight of the fully oxidized sample and m_{red} is the weight of the reduced sample in TGA.

Experimental set-up

A quartz fluidized bed reactor was used to investigate the reactivity of the oxygen carrier with CH₄ in the presence of

SO₂ at CLOU conditions. The experimental set-up simulates the gas–solid contact conditions relevant for a large-scale chemical looping combustion process. Instead of two separate reactors, fuel and air feeds into a single reactor are alternated to create the reduction and oxidation environments in a batch manner. Fig. 1 presents the schematic of the experimental set-up used in this work.

The reactor has a 2 cm inner diameter at the lower section and a 2.5 cm inner diameter section at the top. The increase in diameter at the top section lowers the gas velocity to prevent the smaller particles to leave the reactor. A porous quartz disc inside the reactor supports the particles and allows the gas to flow through the particle bed. A Thermcraft furnace surrounds the reactor to heat it up to the desired temperature. An Omega PX-409 differential pressure transducer measures the pressure fluctuations across the reactor bed and the quartz disc, indicating the fluidization behavior of the particles. For all the experiments, the Cu34Mn66 oxygen carrier inventory was 4 g to ensure there is enough oxygen released for fuel conversion. Experiments were performed at various conditions by changing the temperature and concentrations of CH₄ and SO₂ as shown in the test matrix in Table 1. All the experiments were repeated three times to ensure the reproducibility. Each cycle consisted of a reduction–purge–oxidation pattern. Reduction of the oxygen carrier occurred *via* a reducing mixture of either CH₄–N₂ or CH₄–SO₂–N₂ at a flowrate of 600 ml min^{−1}. Duration of the reduction period was 1 minute for most of the cases. After the reduction cycle, inert N₂ was flowed through the reactor for around 1 minute to purge the reactor, followed by 700 ml min^{−1} air flow to reoxidize the particles. The selected reduction and oxidation flow rates, respectively, correspond to about 2–8 and 2–9 times the theoretical minimum fluidization velocity (U_{mf}). The U_{mf} value is calculated according to the correlations of Kunii and Levenspiel.⁵⁴ Flue gas species (CO, CO₂, CH₄, and O₂) were

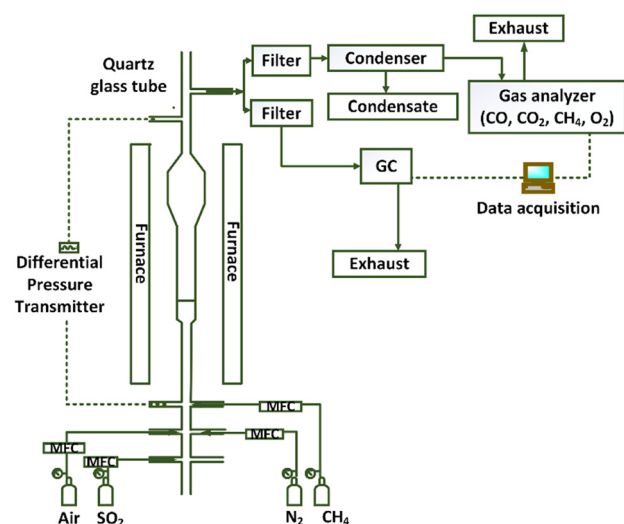


Fig. 1 Schematic of the experimental set-up.

Table 1 Experimental conditions

Case	Oxygen carrier amount (g)	Temperature (°C)	SO ₂ concentration (ppm)	CH ₄ concentration (%)	Reducing time (min)
1	4	900	0, 1000, 3000, 5000	9	1
2	4	900	0, 5000	3, 6, 9	1
3	4	900	5000	9	1, 3, 8
4	4	850, 900, 950	0, 5000	9	1

analyzed at a dry basis by a continuous Rosemount X-stream enhanced analyzer to obtain the reduction profiles for each cycle. H₂ and S species were measured on a wet basis by a gas chromatograph (SRI MG#5 GC); therefore, eliminating the possibility of losing sulfur in condensed water. With a GC, the analysis of each sample takes 8 minutes, which makes it impossible to obtain a continuous profile of SO₂ in a single experiment. Therefore, fresh particles were exposed to different reduction periods in separate experiments to obtain the SO₂ concentration profile against time for a certain cycle. The GC data was collected at the end of each reducing period. So, each data point shown in the concentration profiles presented in the results section represents data from a separate experiment. For the experiments conducted to obtain the SO₂ concentration profiles, the reduction cycle was extended to 8 minutes, instead of 1 min, to see if prolonged exposure to SO₂ results in sulfur accumulation on the oxygen carrier. Additionally, the salt method was employed to qualitatively detect SO₃ in the gas phase as opposed to quantitatively measuring its concentration. This analysis was performed by flowing the fluidized bed reactor effluent from a specific reduction cycle through a small packed-bed reactor containing 1 g of sodium chloride (NaCl) without condensing the water.^{55–57} The temperature of the sample line from the reactor to the packed bed was maintained at 180 °C, while the reactor temperature was maintained at 200 °C. In the presence of H₂O vapor at this temperature, SO₃ reacts with the salt and form sodium sulfate.⁵⁷ The exposed salt sample was titrated by barium perchlorate (Ba(ClO₄)₂) solution to detect the presence of SO₃.⁵⁶

Data evaluation

Reduction performance of the Cu34Mn66 was evaluated using total conversion of CH₄ and yields of CO₂ and CO. Total CH₄ conversion considers formation of both CO and CO₂ as calculated *via* eqn (E2).

$$\text{CH}_4 \text{ conversion (\%)} = \frac{[(X_{\text{CO}} + X_{\text{CO}_2}) / (X_{\text{CO}} + X_{\text{CO}_2} + X_{\text{CH}_4})] \times 100}{\quad} \quad (\text{E2})$$

CO₂ and CO yields were calculated *via* eqn (E3) and (E4).

$$\text{CO}_2 \text{ yield (\%)} = \frac{[X_{\text{CO}_2}] / (X_{\text{CO}} + X_{\text{CO}_2}) \times 100}{\quad} \quad (\text{E3})$$

$$\text{CO yield (\%)} = \frac{[X_{\text{CO}}] / (X_{\text{CO}} + X_{\text{CO}_2}) \times 100}{\quad} \quad (\text{E4})$$

Here, X_i is the molar fraction of the carbon gas species in the reactor outlet stream, with i denoted as CO, CO₂, or unreacted CH₄. The carbon mass balance closure for the 1 minute reduction period is 0.98 ± 0.013 SE. From the sulfur balance around the reactor, the SO₂ loss is calculated by eqn (E5).

$$\text{SO}_2 \text{ loss (mol\%)} = [1 - (X_{\text{SO}_2, \text{out}} \cdot F_{\text{out}}) / (X_{\text{SO}_2, \text{in}} \cdot F_{\text{in}})] \times 100 \quad (\text{E5})$$

Here, $X_{\text{SO}_2, \text{out}}$ is the molar fraction of SO₂ in the flue gas as measured by the GC analyzer and $X_{\text{SO}_2, \text{in}}$ is the inlet molar fraction of SO₂. F_{out} is the calculated outlet molar flow rate, F_{in} is the inlet molar flow rate of the reducing gas. The loss of SO₂ can occur from sulfur accumulation on the Cu34Mn66 particles, SO₃ formation, or both.

Results

Effect of SO₂ on oxygen uncoupling behavior of Cu34Mn66

Fig. 2 depicts the oxygen concentration profile during oxygen uncoupling in N₂ at 900 °C with and without 5000 ppm of SO₂. Oxygen concentration measurement in an empty reactor is also reported to show that the N₂ flow clears the existing air in the reactor and the oxygen detected in presence of the oxygen carrier particles is the oxygen released from the particles. Without SO₂, the oxygen concentration decreases over time, from 6% to 0.6% within 11 minutes. Similar



Fig. 2 O₂ concentration profile during oxygen uncoupling of Cu34Mn66 particles in N₂ at 900 °C with and without 5000 ppm SO₂.

behavior was observed in previous literature.³⁶ In the presence of SO₂, the oxygen concentration is slightly higher until about 2.5%; then, the profile is the same. However, the presence of SO₂ does not affect the total oxygen uncoupling capacity of the particles after 20 cycles of SO₂ exposure. TGA results show that R0 remains at 4 wt% \pm 0.02 wt% for the three cases tested: fresh oxidized particles, particles reduced for 20 cycles with 5000 ppm SO₂, and particles reduced for 20 cycles without SO₂.

Chemical looping combustion of CH₄ without SO₂

Initially, the Cu₃₄Mn₆₆ particles were reduced with a mixture of 9% CH₄-balance N₂ to determine its behavior in the absence of SO₂. Fig. 3 shows the dry basis concentration profiles of the flue gas species with and without SO₂ addition during the 5th cycle at 900 °C. Corresponding CH₄ conversion and yields of CO₂ and CO are plotted in Fig. 4 and 5, respectively. In Fig. 3, during the pre-reduction N₂ purge cycle, O₂ concentration drops, but never reaches zero because the Cu₃₄Mn₆₆ particles start releasing O₂ *via* the CLOU mechanism. Upon CH₄ addition without SO₂, fuel oxidation occurs, producing CO₂ as the primary product and a small amount of CO. Oxidation can occur from two reactions: 1) CH₄ oxidation with the released O₂ in the gas phase and 2) direct reaction with the Cu₃₄Mn₆₆ particles. No H₂ forms during reduction, as confirmed by the GC analysis. As the reduction proceeds, oxygen content of the Cu₃₄Mn₆₆ particles depletes. Consequently, CO₂ concentration decreases while CO and unconverted CH₄ concentrations start to increase. Total CH₄ conversion is between 90–99% throughout the 1 minute reduction period as seen in Fig. 4. Slight amount of CO formation during the first 10 seconds results in a decrease in CO₂ as observed from the CO₂ yield



Fig. 3 Flue gas concentration profiles during the 5th cycle at 900 °C with and without 5000 ppm SO₂. The reducing gas is 9% CH₄-balance N₂ and 9% CH₄-5000 ppm SO₂-balance N₂. The oxidizing gas is air.



Fig. 4 CH₄ conversion as a function of reduction period during the 5th cycle at 900 °C with 9% CH₄-X ppm SO₂-balance N₂ where X: 0, 1000, 3000, 5000.

shown in Fig. 5. It is also worth noting that O₂ release still occurs from the Cu₃₄Mn₆₆ particles (about 1%) during the subsequent inert period after stopping the reducing gas flow of 9% CH₄-balance N₂ as seen in Fig. 3.

Chemical looping combustion of CH₄ with SO₂

Before starting the reduction cycle in the presence of SO₂, a mixture of 9% CH₄-5000 ppm SO₂-balance N₂ was introduced to an empty reactor at 900 °C to evaluate the possibility of homogenous gas-phase reactions between CH₄ and SO₂. No such gas-phase reactions were observed, as confirmed by the



Fig. 5 CO₂ and CO yields as a function of reduction period during the 5th cycle at 900 °C with 9% CH₄-X ppm SO₂-balance N₂ where X: 0, 1000, 3000, 5000.

steady CH_4 concentration profile. Then the reduction cycle was performed in the presence of the oxygen carrier particles with a mixture of 9% CH_4 -X ppm SO_2 -balance N_2 (X: 1000, 3000, 5000 ppm). Fig. 3 depicts the measured flue gas concentration profiles with 5000 ppm SO_2 addition during the 5th cycle at 900 °C. When 5000 ppm SO_2 is added, CO and unreacted CH_4 amounts are higher initially, compared to the SO_2 -free reduction, resulting in a lower production of CO_2 . Increasing the SO_2 concentration from 1000 ppm to 5000 ppm lowers the conversion of CH_4 , while increasing the CO yield resulting in a decrease in CO_2 yield. Concentration profiles in Fig. 3 shows that negligible O_2 release occurs from the particles during the subsequent N_2 purge after reduction, unlike the case without SO_2 , meaning the oxygen content is depleted when SO_2 is present. Based on this observation, the presence of SO_2 may either contribute to particle deactivation or the consumption of available oxygen during the reduction cycle.

An extended reduction cycle was also carried out to observe the impact of SO_2 on CH_4 combustion as a function of the reduction duration up to 9 minutes. Fig. 6 shows the profiles of the flue gas species obtained during complete reduction of Cu34Mn66 with and without 5000 ppm SO_2 . In general, as the duration of the reduction cycle increases, the degree of reduction of the oxygen carrier particles also increases. Without SO_2 , CO_2 concentration gradually decreases while the unreacted CH_4 concentration increases. However, when SO_2 is present, CO_2 and CH_4 concentrations follow a relatively steady profile between ~2.5–4.5 minutes. In addition, the complete reduction of Cu34Mn66 with SO_2 takes slightly longer to complete than the 8 minutes required for SO_2 -free reduction.

Analysis of post-combustion species with GC shows no formation of reduced sulfur gas species such as COS and H_2S

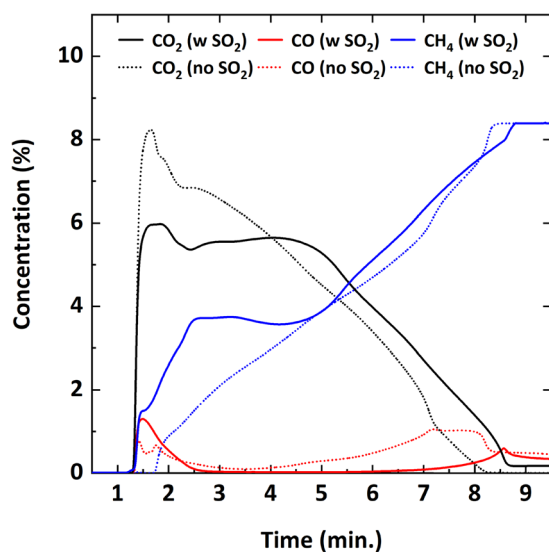


Fig. 6 Flue gas concentration profiles during complete reduction of Cu34Mn66 particles at 900 °C with 9% CH_4 -balance N_2 and 9% CH_4 -5000 ppm SO_2 -balance N_2 .



Fig. 7 Measured SO_2 concentration profile and calculated SO_2 loss (mol%) as a function of reduction period for the 1st, 5th, and 20th cycles at 900 °C with 9% CH_4 -5000 ppm SO_2 -balance N_2 .

during oxygen carrier reduction *via* 9% CH_4 -5000 ppm SO_2 -balance N_2 . Fig. 7 shows measured SO_2 concentration profile along with calculated SO_2 loss as a function of reducing time (up to 8 minutes) for the 1st, 5th, and 20th reduction cycles. For the 8-minute reduction, a single cycle was run due to observed particle agglomeration since the oxygen carrier is completely reduced. For the 1-minute reduction, about 25% of the SO_2 feed is lost throughout and the increasing number of cycles does not affect the SO_2 outlet concentration or the corresponding SO_2 loss during this time. However, for the 3-minute reduction, increasing the cycle number increases the SO_2 loss. When the reduction extent of the particles



Fig. 8 CH_4 conversion as a function of reduction period during the 5th cycle at 900 °C with Y% CH_4 -balance N_2 and Y% CH_4 -5000 ppm SO_2 -balance N_2 where Y: 3, 6, 9.

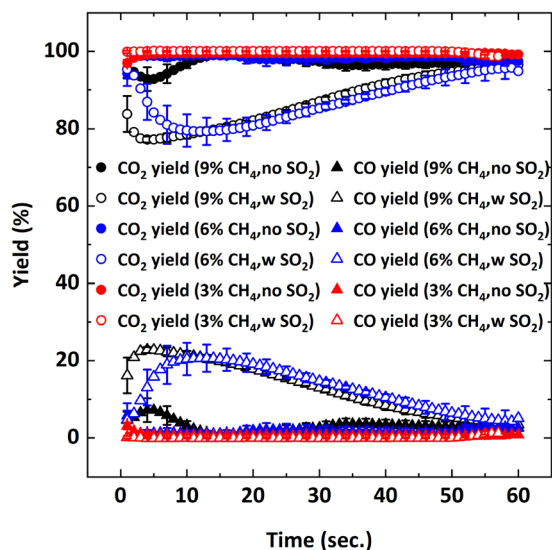


Fig. 9 CO₂ and CO yields as a function of reduction period during the 5th cycle at 900 °C with Y% CH₄-balance N₂ and Y% CH₄-5000 ppm SO₂-balance N₂ where Y: 3, 6, 9.

increases, SO₂ loss also increases with almost no SO₂ left in the flue gas after 8 minutes of reduction.

To determine the likely cause of SO₂ loss in the gas phase, the salt method was employed to detect SO₃ if there is any in the gas phase. SO₂ could possibly be oxidized to SO₃ during the reduction cycle *via* either homogenous gas-phase or surface catalyzed reaction, resulting in the observed loss of SO₂ in the gas-phase. For the 1-minute and 3-minute reduction periods, SO₃ was observed during the 5th and 20th cycles, but not during the 1st cycle. However, for the 8-minute reduction it was observed during the 1st cycle.

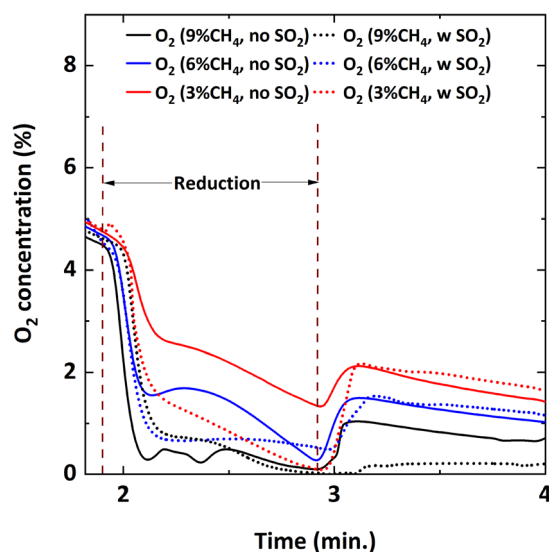


Fig. 10 O₂ concentration profile during the 5th cycle at 900 °C with Y% CH₄-balance N₂ and Y% CH₄-5000 ppm SO₂-balance N₂ where Y: 3, 6, 9.

Since SO₂ is oxidized to SO₃, in order to see if the availability of O₂ changes the effect of SO₂ on the oxygen carrier's reactivity, experiments were run with different CH₄ concentrations in the feed. Oxygen carrier particles were reduced with gas mixtures of Y% CH₄-balance N₂ and Y% CH₄-5000 ppm SO₂-balance N₂ (Y: 3, 6, 9), and the corresponding CH₄ conversions and the yields of CO₂ and CO are provided in Fig. 8 and 9, respectively.

Without SO₂, on average, 95–100% CH₄ conversion occurs and there is not much difference in conversion for all the CH₄ concentrations tested. As discussed before, CH₄ conversion decreases in the presence of SO₂. The adverse outcome of SO₂ is inversely proportional to CH₄ concentration. For example, for 3% CH₄ in the feed, the effect of SO₂ on CH₄ conversion is minimal while the effect is more pronounced for the 9% CH₄ case. To explain this behavior, the oxygen concentration profiles were examined. Fig. 10 shows the O₂ concentration profiles for different CH₄ concentrations with and without 5000 ppm SO₂.

Lowering the CH₄ feed concentration from 9% increases gaseous O₂ availability during the reduction cycle. In other words, for the lower CH₄ concentrations (6% and 3%), more O₂ is released from the oxygen carrier than the amount needed for the combustion reaction with CH₄. However, when SO₂ is added, the availability of O₂ decreases and the O₂ concentration profiles become similar for all the CH₄ concentrations. This could be because the excess oxygen is being used up by SO₂ to become oxidized to SO₃; therefore, indicating a competing oxygen demand between SO₂ and CH₄. However, since Fig. 7 shows that SO₂ loss% increases with the reduction period, it can also be assumed that sulfur interacts with the reduced oxygen carrier. Post combustion analysis of the oxygen carrier is discussed later to investigate this hypothesis.



Fig. 11 CH₄ conversion as a function of reduction period during the 5th cycle with 9% CH₄-balance N₂ and 9% CH₄-5000 ppm SO₂-balance N₂ at different temperatures (850, 900, 950 °C).

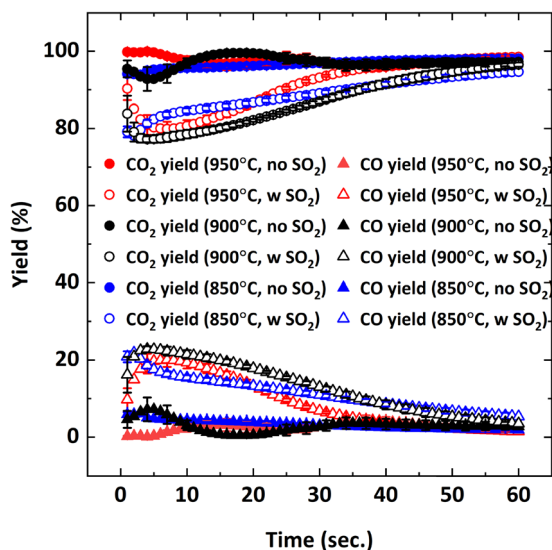


Fig. 12 CO₂ and CO yields as a function of reduction period during the 5th cycle with 9% CH₄-balance N₂ and 9% CH₄-5000 ppm SO₂-balance N₂ at different temperatures (850, 900, 950 °C).

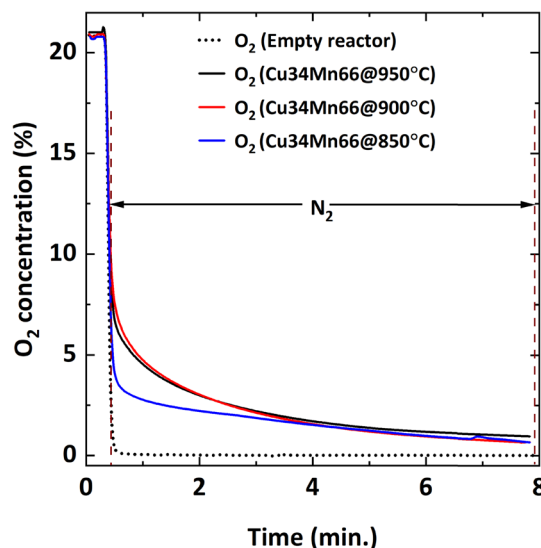


Fig. 13 O₂ concentration profile during oxygen uncoupling of Cu₃₄Mn₆₆ particles at different temperatures (850, 900, 950 °C).

Effect of SO₂ on the combustion of CH₄ at different operating temperatures was also examined since equilibrium O₂ concentration and oxygen release rate of Cu₃₄Mn₆₆ are directly proportional to temperature.³⁶ Fig. 11 illustrates CH₄ conversion and Fig. 12 illustrates the yields of CO₂ and CO at 850, 900 and 950 °C as a function of the reduction period with and without 5000 ppm SO₂ addition.

Fig. 11 shows that decreasing the operating temperature increases the negative effect of SO₂ on CH₄ conversion, with the effect being more pronounced at 850 °C. However, CO₂ and CO yields are similar at different temperatures (Fig. 12), indicating that SO₂ addition increases the unconverted CH₄. To explain the effect of temperature, O₂ uncoupling profiles in N₂ at different temperatures are provided in Fig. 13. The O₂ concentration within the first 2 minutes is lower at 850 °C than those at 900 °C and 950 °C. That explains the lower CH₄ conversion at 850 °C than the conversion at 900 °C and 950 °C without SO₂. When SO₂ is added, the decrease in CH₄ conversion is much larger comparing to the case without SO₂ since there is a competing O₂ demand between SO₂ and CH₄. At 900 and 950 °C, O₂ amount is similar, but CH₄ conversion is higher at 950 °C in the presence of SO₂ while it is similar without SO₂. Therefore, the negative effect of SO₂ on the reactivity of the oxygen carrier with CH₄ becomes more pronounced at lower temperatures.

Effect of SO₂ on multicycle stability of Cu₃₄Mn₆₆

The particles were exposed to 20 reduction-oxidation cycles with 1-minute reduction to determine if the CH₄ conversion degrades after multiple cycles when SO₂ is present. Fig. 14 shows that the concentration profiles of the flue gas species remain consistent throughout 20 cycles, indicating

that the performance of the particles does not degrade over multiple cycles. Following the exposure of oxygen carrier particles to SO₂ over 20 cycles, the particles were reduced without SO₂ with a mixture of 9% CH₄-balance N₂. The purpose of this experiment was to determine if the SO₂-free reduction performance of the SO₂-exposed particles differed from those not initially exposed to SO₂. Fig. 14 shows that during the 21st cycle most of the inlet CH₄ converts to CO₂, which is consistent with the behavior in Fig. 3. So, it can be assumed that the Cu₃₄Mn₆₆ particles recover their reactivity after oxidation and after 20 cycles of SO₂ exposure no permanent deactivation occurs. Additionally, no agglomeration of the particles was observed after 21 cycles.



Fig. 14 Multicycle redox behavior of Cu₃₄Mn₆₆ particles at 900 °C with 9% CH₄-5000 ppm SO₂-balance N₂. Reducing time is 1 minute.

Characterization of Cu34Mn66

The Cu34Mn66 particles tested at various operating conditions were characterized with XRD, XPS, and TPR to evaluate any structural or chemical composition changes following SO₂ exposure. Characterization results of samples reduced with two different gas mixtures, *i.e.*, 9% CH₄-5000 ppm SO₂-balance N₂ and 9% CH₄-balance N₂, at 900 °C are reported here.

Fig. 15 shows the XRD patterns of Cu34Mn66 samples reduced for 1, 3 and 8 minutes for 5 cycles without SO₂ and the fresh oxidized sample for comparison. Powder XRD analysis confirms that fresh, fully oxidized Cu34Mn66 consists of CuMn₂O₄ and Mn₃O₄ crystalline structures. After 1-minute reduction, CuMn₂O₄ converts into major Cu_{1.04}Mn_{0.96}O₂, Mn₃O₄, and minor Cu phases in the absence of SO₂. Extended reduction of Cu34Mn66 for 3 minutes results in additional Cu₂O and MnO crystalline phases, while a fully reduced sample at 8 minutes consists of only Cu and MnO. Fig. 16 shows the XRD patterns for the samples reduced for 1, 3 and 8 minutes for 5 cycles in presence of 5000 ppm SO₂ and the oxidized sample following 1-minute reduction with SO₂. Oxidized sample after 1-minute reduction with 5000 ppm SO₂ has an XRD pattern similar to the pattern of the fresh sample. Reduced samples exposed to SO₂ do not show any sulfate or sulfide presence and the XRD patterns look similar to the patterns with no SO₂ exposure. The only difference is for the sample reduced for 1 minute with SO₂ where the Cu peak is not present. Same thing is observed for the sample reduced for 1 minute for 20 cycles with and without SO₂ (Fig. S1 and S2 in ESI†). An interaction between elemental Cu and SO₂ likely occurs, but the resulting sulfur-copper compound might only be present on the surface and, therefore, not observed in the XRD pattern.



Fig. 15 XRD patterns of samples reduced for 1, 3 and 8 minutes for 5 cycles at 900 °C with 9% CH₄-balance N₂ and a fresh oxidized sample.



Fig. 16 XRD patterns of samples reduced for 1, 3 and 8 minutes for 5 cycles at 900 °C with 9% CH₄-5000 ppm SO₂-balance N₂ and oxidized sample after 1-minute reduction for 5 cycles at 900 °C 9% CH₄-5000 ppm SO₂-balance N₂.

XPS analysis was performed for the samples exposed to SO₂ to determine if any sulfur species form on the surface of the oxygen carrier. Fig. 17 shows the S 2p spectra of 1, 3 and 8 minutes reduced samples with SO₂ exposure collected after the 1st cycle. According to literature, the S 2p_{3/2} peak with a binding energy within 168.2–169.9 eV range indicates sulfate species presence on the surface.^{58–60} S spectra is not observed for the 1-minute reduced sample. However, for 3 and 8 minutes reduced samples, sulfate presence on the surface of the reduced sample is observed. For the 8-minute reduction, there is an additional S 2p_{3/2} position at a binding energy of 161.3 eV. This suggests the formation of sulfide species on a



Fig. 17 S 2p spectra of samples reduced for 1, 3 and 8 minutes for 1 cycle at 900 °C with 9% CH₄-5000 ppm SO₂-balance N₂.

fully-reduced surface.^{58,59,61} When the samples were exposed to SO₂ for multiple redox cycles, the sulfate peak is observed even for the 1-minute reduction after 5 cycles. The S spectra for 5th and 20th cycles are provided in Fig. S3†. For the 20th cycle, the sample was re-oxidized after being exposed to SO₂ for 20 cycles and the sulfate peak in the XPS spectra (Fig. S4†) is not distinctly seen after oxidizing the sample, which explains the behavior in Fig. 14 where the oxygen carrier is regenerated when it is oxidized and the effect of sulfur on the reactivity is not permanent.

Fig. 18 shows Cu 2p spectra of SO₂-exposed samples after 1, 3 and 8 minutes of reduction after the 1st cycle. Deconvolution of spectra reveals three distinct Cu 2p_{3/2} peaks at A (932.2–932.5 eV), B (933.8–934.3 eV), and C (935.7–936.1 eV) positions. On the other hand, in the absence of SO₂, the reduced sample only has A and B peaks (Fig. S5†). From the literature, peak A corresponds to the Cu¹⁺ oxidation and may signify reduced CuMnO₂, Cu, and Cu₂O. Peak B indicates Cu²⁺ oxides, such as CuO and CuMn₂O₄. Finally, peak C coincides with a Cu²⁺ peak that ascribes to the Cu 2p_{3/2} peak of copper sulfate compounds.^{59,62–66} For the SO₂-exposed samples reduced for 3 and 8 minutes, the presence of sulfate peak gets more prominent comparing to 1-minute reduction. For 1-minute reduction, the sulfate peak gets more pronounced as the number of cycles increases to 20 (Fig. S6†).

Fig. 19 shows Mn 2p spectra of reduced samples with SO₂ exposure after different reduction periods. From the literature, Mn 2p_{3/2} peak positions at A (640.4–642.5 eV), B (641.3–641.9 eV), and C (641.6–646.2 eV) correspond to Mn²⁺, Mn³⁺, and Mn⁴⁺ oxidation states, respectively.^{58,67–70} Spectra for all three samples exhibit Mn²⁺, Mn³⁺, and Mn⁴⁺ oxidation states. Since the Mn 2p_{3/2} peak at 642.7 eV observed for MnSO₄ (ref. 71) overlaps with Mn⁴⁺, it is challenging to



Fig. 19 Mn 2p spectra of samples reduced for 1, 3 and 8 minutes for 1 cycle at 900 °C with 9% CH₄-5000 ppm SO₂-balance N₂.

confirm the contribution of Mn towards sulfate formation on the surface.

To see if SO₂ is interacting with the Mn sites, TPR analysis was conducted. Fig. 20 shows the TPR profiles of three samples: fresh oxidized, 1-minute reduced sample with 5000 ppm SO₂ and 1-minute reduced sample without SO₂. Each sample shows two broad overlapping peaks. One is a low-temperature peak between 305–325 °C, and the other is a high-temperature peak between 365–390 °C. The low and high-temperature peaks ascribe to the reduction of Cu²⁺ to Cu, and the reduction of Mn³⁺ and Mn⁴⁺ to Mn²⁺, respectively.^{35,68,72} The low-temperature peak shrinks and shifts to a lower temperature, while the high-temperature



Fig. 18 Cu 2p spectra of samples reduced for 1, 3 and 8 minutes for 1 cycle at 900 °C with 9% CH₄-5000 ppm SO₂-balance N₂.



Fig. 20 H₂-TPR spectra of fresh Cu₃₄Mn₆₆ and samples reduced for 1 minute for 5 cycles at 900 °C with 9% CH₄-balance N₂ and 9% CH₄-5000 ppm SO₂-balance N₂.

peak gets more prominent when the sample is reduced in the presence of SO₂ comparing to the sample reduced without SO₂. Since a clear difference in the reduction profiles for Mn is seen for the samples reduced with and without SO₂, it can be concluded that SO₂ is interacting with Mn in addition to Cu.

Discussion

The results of this study show that SO₂ adversely affects CH₄ conversion and, consequently, the performance of the Cu–Mn oxide in the CLOU process. The possible reasons for the adverse effects of SO₂ are outlined in the following discussion.

According to previous studies, reactions between CH₄ and SO₂ in sub-stoichiometric conditions can form reduced H₂S and COS species^{73–75} *via* heterogeneous catalytic reactions. Both Cu⁷³ and Mn⁷⁶ elements possess catalytic activity for these reactions. Similarly, they can also act as a catalyst for SO₂ oxidation.⁷⁷ Experimental evidence in this study confirms SO₃ presence, but there is no presence of reduced sulfur gaseous species such as H₂S or COS, not even during the extended reduction for 3 and 8 minutes. SO₂ oxidation to SO₃ explains the decrease in O₂ concentration when SO₂ is present during the reduction cycles when CH₄ concentration is less than 9% (Fig. 10). When SO₂ is added, the availability of O₂ decreases because the excess O₂ is being used up by SO₂ to become oxidized to SO₃; therefore, indicating a competing O₂ demand between SO₂ and CH₄.

Based on the XPS results, SO₃ formation is not the only contributor in the reduction of CH₄ conversion with the presence of SO₂. From the XPS results, S 2p and Cu 2p spectra confirm the presence of SO₄^{2–} species after 1-minute reduction, particularly in the form of CuSO₄. Sulfate species might form by the interaction of SO₂ with any of the reduced species such as Cu, CuMnO₂, or Mn₃O₄. The presence of manganese-sulfur surface species remains unconfirmed *via* XPS. Still, the formation of Mn-sulfate species on the surface of the oxygen carrier is possible since previous selective catalytic reduction (SCR) studies show that SO₂ could poison Mn-based catalysts.⁶⁸ In fact, the TPR results shows that SO₂ not only interacts with Cu, but also Mn sites. On the other hand, XRD analysis reveals no sulfate or sulfide crystalline structure formation despite the presence of sulfate on the surface of the oxygen carrier. Therefore, it can be concluded that the crystalline structure of the oxygen carrier is not affected by SO₂.

Decreased CH₄ conversion in the presence of SO₂ could be as a result of the sulfate species on the surface inhibiting O₂ release and/or the heterogeneous CLC reaction *via* lattice oxygen. However, no change in O₂ release was observed in this study (Fig. 2). Therefore, the reduction in CH₄ conversion could be attributed to the heterogeneous CLC reaction being impacted by the presence of sulfate species on the surface of the oxygen carrier, in addition to the

competing oxygen demand between SO₂ and CH₄ using the released O₂ in the gas phase.

Despite the observed decrease in activity, the effect of SO₂ does not worsen over multiple redox cycles. Moreover, the SO₂-exposed sample can be regenerated after oxidation, going back to its original activity before SO₂ exposure.

The extent of reduction of Cu–Mn oxide also affects the influence SO₂ has on the oxygen carrier and, consequently, the fuel conversion. A high degree of reduction produces significant elemental Cu, increasing SO₂ interaction probability, as found during the 3- and 8-minute reduction cycles. For instance, as seen from the XPS analysis, sulfide species form on the surface of the oxygen carrier upon full reduction. Keeping the reduction extent of materials within the CLOU mechanism with a short residence time where there is less Cu might improve the performance of the Cu–Mn oxide. In addition, due to SO₃ formation, the oxygen carrier to fuel ratio might need to be increased to compensate for the oxygen needed for SO₃ formation. It is possible that with a higher oxygen carrier to fuel ratio, the adverse effect of sulfur may not be significant.

Conclusions

The interaction of SO₂ with a bimetallic Cu–Mn oxygen carrier has been investigated under CLOU conditions in a batch fluidized-bed reactor. The results suggest that SO₂ negatively affects total CH₄ conversion and CO₂ yield. However, the oxygen release capacity of the Cu–Mn oxide remains unaffected. The reactivity with CH₄ decreases as the SO₂ concentration increases. Increasing the temperature from 850 to 950 °C improves the conversion of CH₄. During the reduction cycle, SO₂ oxidizes to SO₃, creating a competing oxygen demand in the reactor with the fuel, and lowers CH₄ conversion if there is not enough oxygen for both reactions. Hence, decreasing the feed CH₄ concentration with the same oxygen carrier amount lowers the impact of SO₂ on CH₄ conversion.

Additionally, characterization of SO₂-exposed oxygen carriers *via* XPS shows that sulfate species form on the surface, possibly causing further reduction in CH₄ conversion. However, the presence of sulfate does not cause a permanent deactivation of the particles. The reduction in CH₄ conversion does not worsen during 20 redox cycles with 5000 ppm SO₂ and the particles can be fully regenerated during the oxidation cycle. Nevertheless, further studies are required in a continuous CLOU unit to determine the sulfur effect during a long-time operation.

Author contributions

Turna Barua: investigation, methodology, formal analysis, writing – original draft. Bihter Padak: conceptualization, funding acquisition, methodology, project administration, supervision, writing – review & editing.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors are grateful for the financial support provided by the National Science Foundation's (NSF) Combustion and Fire Systems Program (award number 1606874). The authors would also like to thank Madeline Talebi for her assistance with conducting some of the experiments and synthesis of some of the oxygen carriers used in the experiments. She has also helped editing the language of the paper.

References

- Monthly Energy Review: Electricity, U.S Energy Information Administration, 2021.
- M. Bui, C. S. Adjiman, A. Bardow, E. J. Anthony, A. Boston, S. Brown, P. S. Fennell, S. Fuss, A. Galindo, L. A. Hackett, J. P. Hallett, H. J. Herzog, G. Jackson, J. Kemper, S. Krevor, G. C. Maitland, M. Matuszewski, I. S. Metcalfe, C. Petit, G. Puxty, J. Reimer, D. M. Reiner, E. S. Rubin, S. A. Scott, N. Shah, B. Smit, J. P. M. Trusler, P. Webley, J. Wilcox and N. Mac Dowell, *Energy Environ. Sci.*, 2018, **11**, 1062–1176.
- G. De Guido, M. Compagnoni, L. A. Pellegrini and I. Rossetti, *Front. Chem. Sci. Eng.*, 2018, **12**, 315–325.
- J. Gibbins and H. Chalmers, *Energy Policy*, 2008, **36**, 4317–4322.
- D. Y. C. Leung, G. Caramanna and M. M. Maroto-Valer, *Renewable Sustainable Energy Rev.*, 2014, **39**, 426–443.
- J. Adánez, A. Abad, T. Mendiara, P. Gayán, L. F. de Diego and F. García-Labiano, *Prog. Energy Combust. Sci.*, 2018, **65**, 6–66.
- I. Adánez Rubio, P. Gayán, A. Abad, F. García-Labiano, L. de Diego and J. Adánez, CO₂ Capture in Coal Combustion by Chemical-Looping with Oxygen Uncoupling (CLOU) with a Cu-based Oxygen-Carrier, 2011.
- M. Anheden, A. S. Nasholm and G. Svedberg, *Chemical-looping combustion - Efficient conversion of chemical energy in fuels into work*, Amer Soc Mechanical Engineers, New York, 1995.
- M. M. Hossain and H. I. de Lasa, *Chem. Eng. Sci.*, 2008, **63**, 4433–4451.
- M. Ishida and H. G. Jin, *Ind. Eng. Chem. Res.*, 1996, **35**, 2469–2472.
- J. Li, H. Zhang, Z. Gao, J. Fu, W. Ao and J. Dai, *Energy Fuels*, 2017, **31**, 3475–3524.
- A. Lyngfelt, B. Leckner and T. Mattisson, *Chem. Eng. Sci.*, 2001, **56**, 3101–3113.
- I. Adánez-Rubio, A. Abad, P. Gayán, L. F. de Diego, F. García-Labiano and J. Adánez, *Int. J. Greenhouse Gas Control*, 2013, **12**, 430–440.
- P. Gayan, I. Adanez-Rubio, A. Abad, L. F. de Diego, F. Garcia-Labiano and J. Adanez, *Fuel*, 2012, **96**, 226–238.
- T. Mattisson, H. Leion and A. Lyngfelt, *Fuel*, 2009, **88**, 683–690.
- Z. Zhou, L. Han, O. Nordness and G. M. Bollas, *Appl. Catal., B*, 2015, **166–167**, 132–144.
- M. Arjmand, H. Leion, A. Lyngfelt and T. Mattisson, *Int. J. Greenhouse Gas Control*, 2012, **8**, 56–60.
- P. Hallberg, D. Z. Jing, M. Ryden, T. Mattisson and A. Lyngfelt, *Energy Fuels*, 2013, **27**, 1473–1481.
- C. Linderholm, A. Lyngfelt, A. Cuadrat and E. Jerndal, *Fuel*, 2012, **102**, 808–822.
- J. S. Dennis, C. R. Muller and S. A. Scott, *Fuel*, 2010, **89**, 2353–2364.
- Ø. Langørgen, I. Saanum and N. E. L. Haugen, *Energy Procedia*, 2017, **114**, 352–360.
- M. M. Hossain and H. I. de Lasa, *Chem. Eng. Sci.*, 2008, **63**, 4433–4451.
- T. Mattisson, A. Lyngfelt and H. Leion, *Int. J. Greenhouse Gas Control*, 2009, **3**, 11–19.
- I. Adánez-Rubio, M. Arjmand, H. Leion, P. Gayán, A. Abad, T. Mattisson and A. Lyngfelt, *Energy Fuels*, 2013, **27**, 3918–3927.
- G. Azimi, H. Leion, M. Rydén, T. Mattisson and A. Lyngfelt, *Energy Fuels*, 2012, **27**, 367–377.
- M. Rydén, A. Lyngfelt and T. Mattisson, *Energy Procedia*, 2011, **4**, 341–348.
- M. Rydén, A. Lyngfelt and T. Mattisson, *Int. J. Greenhouse Gas Control*, 2011, **5**, 356–366.
- N. Mohammad Pour, H. Leion, M. Rydén and T. Mattisson, *Energy Fuels*, 2013, **27**, 6031–6039.
- P. Mungse, G. Saravanan, T. Uchiyama, M. Nishibori, Y. Teraoka, S. Rayalu and N. Labhsetwar, *Phys. Chem. Chem. Phys.*, 2014, **16**, 19634–19642.
- A.-M. Azad, A. Hedayati, M. Rydén, H. Leion and T. Mattisson, *Energy Technol.*, 2013, **1**, 59–69.
- F. C. M. Driessens and G. D. Rieck, *Z. Anorg. Allg. Chem.*, 1967, **351**, 48–62.
- I. Adánez-Rubio, A. Abad, P. Gayán, L. F. de Diego and J. Adánez, *Fuel*, 2018, **212**, 605–612.
- I. Adánez-Rubio, A. Abad, P. Gayán, F. García-Labiano, L. F. de Diego and J. Adánez, *Appl. Energy*, 2017, **208**, 561–570.
- C. F. Patzschke, M. E. Boot-Handford, Q. Song and P. S. Fennell, *Chem. Eng. J.*, 2021, **407**, 127093.
- D. Hosseini, Q. Imtiaz, P. M. Abdala, S. Yoon, A. M. Kierzkowska, A. Weidenkaff and C. R. Muller, *J. Mater. Chem. A*, 2015, **3**, 10545–10550.
- I. Adánez-Rubio, M. T. Izquierdo, A. Abad, P. Gayán, L. F. de Diego and J. Adánez, *Int. J. Greenhouse Gas Control*, 2017, **65**, 76–85.
- I. Adánez-Rubio, A. Abad, P. Gayán, I. Adánez, L. F. de Diego, F. García-Labiano and J. Adánez, *Energy Fuels*, 2016, **30**, 5953–5963.
- I. Adánez-Rubio, A. Abad, P. Gayán, L. F. de Diego and J. Adánez, *Fuel*, 2018, **212**, 605–612.
- R. F. Pachler, K. Mayer, S. Penthor, M. Kollerits and H. Hofbauer, *Int. J. Greenhouse Gas Control*, 2018, **71**, 86–94.
- E. Ksepko, R. V. Siriwardane, H. Tian, T. Simonyi and M. Sciazko, *Energy Fuels*, 2012, **26**, 2461–2472.
- P. Adolphi, M. Störr, P. G. Mahlberg, H. H. Murray and E. M. Ripley, *Int. J. Coal Geol.*, 1990, **16**, 185–188.

- 42 J. H. Brady and J. R. Holum, *Fundamentals of Chemistry*, John Wiley & Sons, 1988.
- 43 P. D. Maier and H. P. Dibbs, *Thermochim. Acta*, 1974, **8**, 155.
- 44 I. Adánez-Rubio, A. Abad, P. Gayán, F. García-Labiano, L. F. de Diego and J. Adánez, *Appl. Energy*, 2014, **113**, 1855–1862.
- 45 R. Pérez-Vega, I. Adánez-Rubio, P. Gayán, M. T. Izquierdo, A. Abad, F. García-Labiano, L. F. de Diego and J. Adánez, *Int. J. Greenhouse Gas Control*, 2016, **46**, 28–38.
- 46 M. Arjmand, M. Rydén, H. Leion, T. Mattisson and A. Lyngfelt, *Ind. Eng. Chem. Res.*, 2014, **53**, 19488–19497.
- 47 K. Wang, X. Tian and H. Zhao, *Appl. Energy*, 2016, **166**, 84–95.
- 48 C. Chung, Y. Pottimurthy, M. Xu, T.-L. Hsieh, D. Xu, Y. Zhang, Y.-Y. Chen, P. He, M. Pickarts, L.-S. Fan and A. Tong, *Appl. Energy*, 2017, **208**, 678–690.
- 49 B. Wang, R. Yan, D. H. Lee, D. T. Liang, Y. Zheng, H. Zhao and C. Zheng, *Energy Fuels*, 2008, **22**, 1012–1020.
- 50 C. R. Forero, P. Gayán, F. García-Labiano, L. F. de Diego, A. Abad and J. Adánez, *Int. J. Greenhouse Gas Control*, 2010, **4**, 762–770.
- 51 L. F. de Diego, F. García-Labiano, P. Gayán, A. Abad, A. Cabello, J. Adánez and G. Sprachmann, *Int. J. Greenhouse Gas Control*, 2014, **28**, 168–179.
- 52 X. Tian, K. Wang, H. Zhao and M. Su, *Proc. Combust. Inst.*, 2017, **36**, 3381–3388.
- 53 L. F. de Diego, P. Gayán, F. García-Labiano, J. Celaya, A. Abad and J. Adánez, *Energy Fuels*, 2005, **19**, 1850–1856.
- 54 D. Kunii and O. Levenspiel, *Fluidization Engineering*, Butterworth-Heinemann, Boston, 1991.
- 55 E. Vainio, D. Fleig, A. Brink, K. Andersson, F. Johnsson and M. Hupa, *Energy Fuels*, 2013, **27**, 2767–2775.
- 56 N. N. Choudhury and B. Padak, *Int. J. Greenhouse Gas Control*, 2016, **51**, 165–175.
- 57 D. Fleig, E. Vainio, K. Andersson, A. Brink, F. Johnsson and M. Hupa, *Energy Fuels*, 2012, **26**, 5537–5549.
- 58 A. V. Naumkin, A. Kraut-Vass, S. W. Gaarenstroom and C. J. Powell, *NIST X-ray Photoelectron Spectroscopy Database, NIST Standard Reference Database Number 20*, National Institute of Standards and Technology, Gaithersburg MD, 2000, p. 20899, DOI: [10.18434/T4T88K](https://doi.org/10.18434/T4T88K).
- 59 A. Galtayries and J.-P. Bonnelle, *Surf. Interface Anal.*, 1995, **23**, 171–179.
- 60 M. Fantauzzi, D. Atzei, B. Elsener, P. Lattanzi and A. Rossi, *Surf. Interface Anal.*, 2006, **38**, 922–930.
- 61 M. Kundu, T. Hasegawa, K. Terabe, K. Yamamoto and M. Aono, *Sci. Technol. Adv. Mater.*, 2008, **9**, 035011.
- 62 V. Hayez, A. Franquet, A. Hubin and H. Terryn, *Surf. Interface Anal.*, 2004, **36**, 876–879.
- 63 H.-Y. Chen, Y.-C. Lin and J.-S. Lee, *Appl. Surf. Sci.*, 2015, **338**, 113–119.
- 64 D. Fang, J. Xie, D. Mei, Y. Zhang, F. He, X. Liu and Y. Li, *RSC Adv.*, 2014, **4**, 25540.
- 65 V. A. M. Brabers and F. V. Setten, *J. Phys. D: Appl. Phys.*, 1983, **16**, L169–L172.
- 66 K. K. Shukla, P. Shahi, G. S. A. Kumar, A. K. Ghosh, R. Singh, N. Sharma, A. Das, A. K. Sinha, A. G. Joshi, A. K. Nigam and S. Chatterjee, *RSC Adv.*, 2015, **5**, 83504–83511.
- 67 H. W. Nesbitt and D. Banerjee, *Am. Mineral.*, 1998, **83**, 305–315.
- 68 S. Xiong, Y. Peng, D. Wang, N. Huang, Q. Zhang, S. Yang, J. Chen and J. Li, *Chem. Eng. J.*, 2020, **387**, 124090.
- 69 J. S. Foord, R. B. Jackman and G. C. Allen, *Philos. Mag. A*, 2006, **49**, 657–663.
- 70 M. C. Biesinger, B. P. Payne, A. P. Grosvenor, L. W. M. Lau, A. R. Gerson and R. S. C. Smart, *Appl. Surf. Sci.*, 2011, **257**, 2717–2730.
- 71 B. R. Strohmeier and D. M. Hercules, *J. Phys. Chem.*, 1984, **88**, 4922–4929.
- 72 P. Mungse, G. Saravanan, T. Uchiyama, M. Nishibori, Y. Teraoka, S. Rayalu and N. Labhsetwar, *Phys. Chem. Chem. Phys.*, 2014, **16**, 19634–19642.
- 73 S. E. Mousavi, H. Pahlavanzadeh, M. Khani, H. Ale Ebrahim and A. Mozaffari, *Iran. J. Chem. Chem. Eng.*, 2018, **15**, 94–107.
- 74 J. Sarlis and D. Berk, *Chem. Eng. Commun.*, 1995, **140**, 73–85.
- 75 A. S. Bobrin, V. I. Anikeev, A. Yermakova and V. A. Kirillov, *React. Kinet. Catal. Lett.*, 1989, **40**, 363–367.
- 76 N. Shikina, S. Khairulin, S. A. Yashnik, T. N. Teryaeva and Z. Ismagilov, *Eurasian Chem.-Technol. J.*, 2015, **17**, 129.
- 77 F. Rabiee and K. Mahanpoor, *Int. J. Ind. Chem.*, 2018, **9**, 27–38.