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Charge–orbital synergistic engineering of $\text{TM@Ti}_3\text{C}_2\text{O}_{1-x}\text{B}_x$ for highly selective CO_2 electrochemical reduction†

Jiahe Peng,^{ab} Zuhao Shi,^{ib ab} Jizhou Jiang,^{ib c} Peng Zhang,^{ib d} Jyh-Ping Hsu^{ib e} and Neng Li^{ib *abc}

Inspired by MXene nanosheets and their regulation of surface functional groups, a series of Ti_3C_2 -MXene-based single TM atom electrocatalysts with a doped boron (B) atom ($\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$, TM is V, Cr, Mn, Fe, Co or Ni, $x = 0.11$) are proposed for achieving a high performance catalytic CO_2 reduction reaction (CO_2RR). The results reveal that the doped B atom involves in the adsorption reaction of CO_2 molecules and CO intermediates in the CO_2RR . The TM-to-C and B-to-C π -back bonding contribute to the activation of the CO_2 molecules and CO intermediates in the CO_2RR . Enough electrons from the single TM atom and B atom occupied orbitals can be injected into the CO_2 molecules and *CO intermediates through direct bonding interactions, which effectively alleviates the difficulty of the first hydrogenation reaction step and further helps CO reduction towards CH_4 . The calculated values of ΔG for the first hydrogenation reaction and the formation of *CHO on $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ are significantly smaller than those of other single-atom catalysts (SACs). $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ is found to have the highest electrocatalytic activity with a limiting potential of ~ 0.40 V and exhibits a high selectivity for obtaining CH_4 through the CO_2RR compared with the hydrogen evolution reaction. This work is expected to open a research path for engineering the charge–orbital state of the innate atoms of a substrate based on mechanistic insights, which guides the rational design of highly selective MXene-based CO_2RR electrocatalysts.

1. Introduction

The rapid economic development and urbanization have already caused a sharp increase in the global concentration of carbon

New concepts

Although the development of efficient atomic electrocatalysts to resolve activity and selectivity issues of the CO_2 reduction reaction (CO_2RR) continues to receive increasing attention, it is still challenging. Most of the previous works in this field have primarily focused on TM atoms and introducing a nonmetal atom to modulate the interplay between the TM active centers and multiple intermediates. Herein, we propose a novel mechanism of introducing a boron element, which takes the advantage of charge–orbital synergism to activate the TM atoms of dual atomic electrocatalysts for the CO_2RR . The charge and orbital order controlling resulting from the combined effect of the single TM atoms and B atom can effectively activate CO_2 molecules and CO intermediates in the CO_2RR process. The TM-to-C, and B-to-C π -back bonding contribute to the activation of CO_2 molecules and CO intermediates in the CO_2RR . Based on this mechanism, we further design a series of TM atoms supported by a monolayer B-doped $\text{Ti}_3\text{C}_2\text{O}_2$ ($\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$) for the CO_2RR . Following a comprehensive screening strategy, $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ are identified as promising catalysts with high activity and selectivity. Moreover, we identify the valence states of TM as a valid descriptor to evaluate the CO_2RR performance on such catalysts, realizing a fast pre-screening of other potential dopants on the MXene-based CO_2RR SACs. Our work thus provides a new direction for the rational design of MXene-based dual atomic electrocatalysts.

dioxide (CO_2).^{1,2} The global warming resulting from the excessive consumption of fossil fuels not only poses a serious threat to global economic security but also affects resources, energy, ecology, and food security, among others. It also poses a severe challenge to human survival.^{3,4} In an attempt to alleviate the increase in the average global temperature, Paris Climate Change Conference signed the Paris Agreement.^{5–7}

^a State Key Laboratory of Silicate Materials for Architectures, Wuhan University of Technology, Wuhan 430070, China. E-mail: lineng@whut.edu.cn

^b Shenzhen Research Institute of Wuhan University of Technology, Shenzhen 518000, Guangdong, China

^c School of Chemistry and Environmental Engineering, School of Environmental Ecology and Biological Engineering, Novel Catalytic Materials of Hubei Engineering Research Center, Wuhan Institute of Technology, Wuhan, 430205, China

^d State Center for International Cooperation on Designer Low-Carbon & Environmental Materials (CDLCEM), School of Materials Science and Engineering, Zhengzhou University, Zhengzhou 450001, Henan, China

^e Department of Chemical Engineering, National Taiwan University, Taipei, 10617, China

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The electrochemical reduction of CO₂ to produce carbon-based fuels and chemicals has great potential in solving current environmental problems and benefiting the economy of our society.^{8,9} However, the high chemical stability of CO₂ making it hard to be activated at high conversion rates and selectivity is a crucial issue.^{10,11} In addition, the CO₂RR pathway includes a chemical dehydration step leading to adsorption of CO,^{12–14} which may act as a poison blocking the surface and impeding further protonation steps because of the high desorption energy and a further reduction barrier.^{15–17}

In recent years, intense research has been carried out to discover efficient, selective and stable cathode catalysts. Copper-based materials have been widely regarded as efficient cathode materials for the CO₂RR, producing diverse products, such as carbon monoxide (CO), formic acid (HCOOH), methanol (CH₃OH), methane (CH₄), ethanol (C₂H₅OH), and ethylene (C₂H₄).^{18–20} Intense optimization of the catalyst materials and reaction conditions has resulted in significant progress over the past few years.^{21,22} Irabien *et al.* found that the Cu^{II} atoms in the HKUST-1 metal–organic framework ([Cu₃(μ₆-C₉H₃O₆)₂]_n) can promote the electrocatalytic conversion of CO₂ to alcohols.²³ The subsequent research studies have reported Cu-based bimetallic electrocatalysts through doping other transition metal atoms to further improve the CO₂ conversion performance.^{24,25} Therefore, downsizing the metal nanostructures to singly dispersed metal atoms is highly desirable for maximizing the efficiency of catalytically active metal sites. Recently, single-atom catalysts (SACs) containing isolated metal atoms dispersed on solid substrates have attracted significant attention in the field of catalysis.

Single transition metal (TM) atoms anchored on a substrate are widely used as the active sites in the CO₂RR.^{26,27} The excellent performance of transition metal-based catalysts can be ascribed to their occupied d orbitals, which can donate electrons to the antibonding orbitals of gas molecules.^{28–30} Being a nonmetal element, the boron (B) atom in a molecular catalyst with sp² or sp³ hybridization has exhibited great potential for N₂ fixation.^{31–33} The B atom also shows the ability

to activate CO₂ molecules in the CO₂RR.^{34,35} Tang *et al.* found that the bare edge B on armchair boron nitride nanoribbons is a promising candidate for the catalysis of the CO₂RR.³⁶ In light of the above discussion, the active center of catalysts could be either a single TM atom coordinated by non-metal atoms or a B center. It can be inferred that if B serves as the coordination atom of TM, the active sites with the presence of both B and TM may provide stronger activation capacity.

The results of our previous study indicated that the combined effect of TM and B can effectively and selectively ameliorate the performance of electrochemical N₂ reduction.³⁷ Being an isoelectronic body of N₂, CO is likely to be activated owing to the combined effect of TM and B atoms. Liu *et al.* successfully fabricated single atomic Fe sites anchored on B and N co-doped carbon (Fe-SA/BNC), which showed excellent CO₂RR activity, achieving the highest faradaic efficiency of ~94% at –0.7 V compared with RHE.³⁸ However, there is a lack of an in-depth discussion on the mechanisms of CO₂ activation. A detailed understanding of these mechanisms at the atomic scale is of great importance for clarifying the synergistic effect between TM and B atoms that contributes to the CO₂RR. This is beneficial to developing a suitable catalyst with a high catalytic selectivity for the CO₂RR.

In addition to the active center, the choice of substrate is another significant factor that can influence SACs' performance.^{39,40} MXenes (*e.g.*, Ti₃C₂T_x, Ti₂CT_x and Mo₂CT_x), having a general formula of M_{n+1}X_nT_x (*n* = 1 to 4) with M, X, and T_x being the early transition metals, C or N, and the surface functional groups (usually as –O, –OH or –F), respectively,^{41,42} have been considered as potential substrates for various applications,^{43,44} especially in the field of SACs.^{45–47} Compared with other supports (*e.g.*, doped graphene, MoS₂, and g-C₃N₄), MXenes have not only the highest metallic conductivity but also abundant surface functional groups for anchoring firmly single atoms.⁴⁸ Recently, MXenes have been widely used as the substrates of SACs for the CO₂RR.⁴⁹ Zhao *et al.* have successfully demonstrated that a Pt@Ti₃C₂T_x-based SAC could readily capture CO₂ or aniline, yielding value-added amides with a high conversion and selectivity.⁵⁰ The subsequent study of Yang *et al.* has produced single atom Cu-immobilized Ti₃C₂T_x for an electrocatalytic CO₂RR to produce methanol with a high faradaic efficiency of 59.1%.⁵¹ However, relevant theoretical analyses of the role of B in the MXene-based SACs' CO₂RR are very limited. The SACs of single atoms fixed on the Ti₃C₂T_x mixed functional group surface have been successfully synthesized experimentally.^{52,53}

In particular, the B atom has been successfully introduced into Ti₃C₂T_x to replace the O functional group on its surface⁵⁴ and the single Ru atom anchored on the B-doped Ti₃C₂T_x exhibited outstanding performance for the hydrogen evolution reaction (HER).⁵⁵ Relevant reported results provide valuable information for studying the synergistic effect of TM and B atoms on the CO₂RR performance of Ti₃C₂T_x-based SACs. Our previous theoretical analysis also showed that the CO₂RR performance of MXene-based SACs can be efficiently improved through controlling of functional groups.⁵⁶ It suggests that an in-depth study on the performance of the CO₂RR through



Neng Li

Congratulations on the 10th anniversary and best wishes to Materials Horizons. Our first paper was published in Materials Horizons, and then chosen as the outstanding paper in the in 2018, and selected as "highlighting the members of the Materials Horizons Advisory Board" in 2021. The best way to express our gratitude is to continue to interact closely with the community, and report our significant research advances, such as novel materials, and new reaction

mechanisms. We are proud to have solid and excellent cooperation with such an esteemed journal. Here's to another decade, we would like to further contribute this esteemed journal.

controlling of MXene-based SACs with a mixed functional group surface is necessary. Therefore, we discuss in this study the feasibility of anchoring low cost 3d single TM (TM is V, Cr, Mn, Fe, Co or Ni) atoms on the $\text{Ti}_3\text{C}_2\text{T}_x$'s surface groups as a mixture of B and O ($\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$) to explore the synergistic effect of TM and B atoms on $\text{Ti}_3\text{C}_2\text{T}_x$ -based SACs for the CO_2RR . The analysis is based on DFT calculations.

As expected, the results of our computation show that the combined effect of single TM atoms and B atom markedly activated CO_2 molecules and CO intermediates after adsorption. This can be attributed to the injection of adequate charge from the occupied p and d orbitals of B and TM, respectively. In addition, the introduction of the B atom into surface groups can significantly enhance the single atom anchoring capacity. The adsorbed CO_2 can be further reduced to CH_4 on the $\text{Fe}@ \text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ SAC with a rather low overpotential of 0.40 eV. The stability of $\text{TM}@ \text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ is systematically evaluated, and the results obtained reveal that the designed catalyst can be synthesized promisingly.

2. Computational details

All calculations were performed on the Vienna *ab initio* simulation package (VASP).⁵⁷ The Perdew–Burke–Ernzerhof (PBE) functional in the generalized gradient approximation (GGA) was used to express electron exchange correlation⁵⁸ and a projector augmented wave (PAW) method was applied to describe the pseudo-potentials.⁵⁹ Dispersion-corrected DFT-D3 schemes were employed to describe possible van der Waals (VDW) interactions.⁶⁰ The energy cut-off for the plane-wave basis was set at 450 eV. All configurations were based on a $3 \times 3 \times 1$ supercell of $\text{Ti}_3\text{C}_2\text{T}_x$ with a vacuum layer larger than 20 Å in the z direction to prevent the interaction between periodic images. The Brillouin zone in reciprocal space was sampled using the Monkhorst–Pack scheme with $5 \times 5 \times 1$ k-point grids for geometry optimization and electronic structure calculations.

All structures were fully relaxed until the forces became smaller than $0.01 \text{ eV } \text{Å}^{-1}$; the convergence threshold in electronic relaxation was set at 10^{-5} eV using the conjugate gradient algorithm. The atomic charge was calculated through Bader's charge population analysis.⁶¹ The atomic structures were analyzed using the VESTA code.⁶² The effect of water was considered using the VASPsol code.^{63–65} The stability of catalysts was evaluated using *ab initio* molecular dynamics (AIMD) simulations carried out by placing MXene in a vacuum or liquid water and lasted for 10.0 ps with a time step of 1 fs in the NVT ensemble.

The binding energy (E_b) and cohesive energy (E_c) of the transition metals on $\text{Ti}_3\text{C}_2\text{T}_x$ can be evaluated using

$$E_b = E_{\text{TM-MX}} - E_{\text{MX}} - E_{\text{TM}} \quad (1)$$

$$E_c = E_{(\text{TMbulk})/N} - E_{\text{TM}}, \quad (2)$$

where $E_{\text{TM-MX}}$, E_{MX} , and E_{TM} are the total energies of $\text{Ti}_3\text{C}_2\text{T}_x$ with and without transition metals and the isolated transition metals, respectively.⁶⁶ $E_{(\text{TMbulk})}$ and N are the energy of the bulk

crystal unit cell of the corresponding transition metals and the number of atoms in the unit cell, respectively. The adsorption energy (ΔE_{ad}) of the different intermediates can be calculated using

$$\Delta E_{\text{ad}} = E_{\text{X}^*} - E^* - E_{\text{X}}, \quad (3)$$

where E_{X^*} , E^* , and E_{X} represent the energies of the total system after adsorption, the catalyst, and the adsorbates, respectively.⁶⁷

A computational hydrogen electrode (CHE) model proposed by Nørskov *et al.*⁶⁸ was used to calculate the Gibbs free energy (G) of the reaction for the CO_2RR elementary steps involving ($\text{H}^+ + \text{e}^-$) pair transfer. In the CHE model, $\text{H}^+ + \text{e}^- \rightleftharpoons 1/2\text{H}_2$ (g) is in equilibrium with gaseous H_2 at 0 V vs. the reversible hydrogen electrode (RHE) for all pH values at 1 bar of H_2 pressure. The Gibbs free energy of each elementary step was calculated at 298.15 K using the following eqn (4):

$$G = E + \text{ZPE} + \int C_p dT - TS + G_{\text{pH}} \quad (4)$$

where E is the total energy and T is the temperature. ZPE, $\int C_p dT$, and TS are the zero-point energy, the enthalpy changes from 0 to T K contributed by molecular vibrations, and the entropy correction based on the calculated vibrational frequencies, respectively. G_{pH} represents the free energy correction for pH, defined as $G_{\text{pH}} = -k_B T \ln[\text{H}^+] = k_B T \ln 10 \times \text{pH}$, where k_B is the Boltzmann constant.

In this work, the pH value was set at 0 for strongly acidic conditions.⁶⁸ The dependence of the potential (U) on the RHE required for each protonation step is expressed as $U = -\Delta G/e$, where ΔG is the free energy change along each CO_2RR pathway. The limiting potential (U_L) was obtained from the maximum $\Delta G(\Delta G_{\text{max}})$ of each elementary step. Since the Gibbs free energy of the O_2 , CO, and CO_2 molecules in a gas-phase species cannot be calculated precisely using conventional PBE exchange–correlation function,^{14,69} the free energy of the experimental thermodynamic data was used as a ref. 70 and 71. The free energy of liquid-phase molecules (H_2O and CH_3OH) was calculated by considering them as an ideal gas and using their corresponding vapor fugacity (f) at $\mu_{(\text{liquid})} = \mu_{(\text{vapor})}$.¹⁴

3. Results and discussion

3.1. Stability of single atoms embedded on $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$

The diffusion and aggregation of the dispersed single atoms on a support can lower the activity and durability of SACs, whereas $\text{Ti}_3\text{C}_2\text{T}_x$ can capture and stabilize the heteroatom effectively through creating tight bonds between the TM atoms and surface groups. In general, the hollow position consisting of three oxygen groups is the main site on pure $\text{Ti}_3\text{C}_2\text{T}_x$ to fix single TM atoms.^{72,73} Therefore, we established a pure $\text{Ti}_3\text{C}_2\text{T}_x$ with oxygen groups covering its surface ($\text{Ti}_3\text{C}_2\text{O}_2$). As introduced previously, the B atom can replace the position of the surface oxygen functional group and be used as the coordination atom to fix the TM atom when doped in $\text{Ti}_3\text{C}_2\text{O}_2$ ($\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$). As shown in Fig. 1a and b, two possible hollow sites (H1 and H2) might present to locate the TM on the surface of $\text{Ti}_3\text{C}_2\text{O}_2$

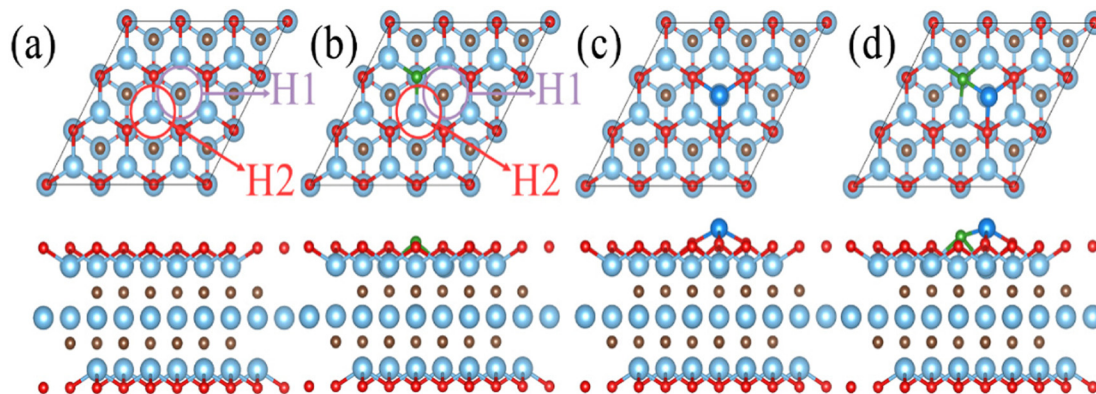


Fig. 1 Optimized models of (a) pure $\text{Ti}_3\text{C}_2\text{O}_2$ and (b) $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$; color scheme: Ti, cyan; C, brown; O, red; B, green; and TM, blue. Circles represent the possible sites of TM on $\text{Ti}_3\text{C}_2\text{O}_2$. Optimized structures of (c) pure $\text{Ti}_3\text{C}_2\text{O}_2$ and (d) $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ monolayer confining TM single atoms at the H1 site.

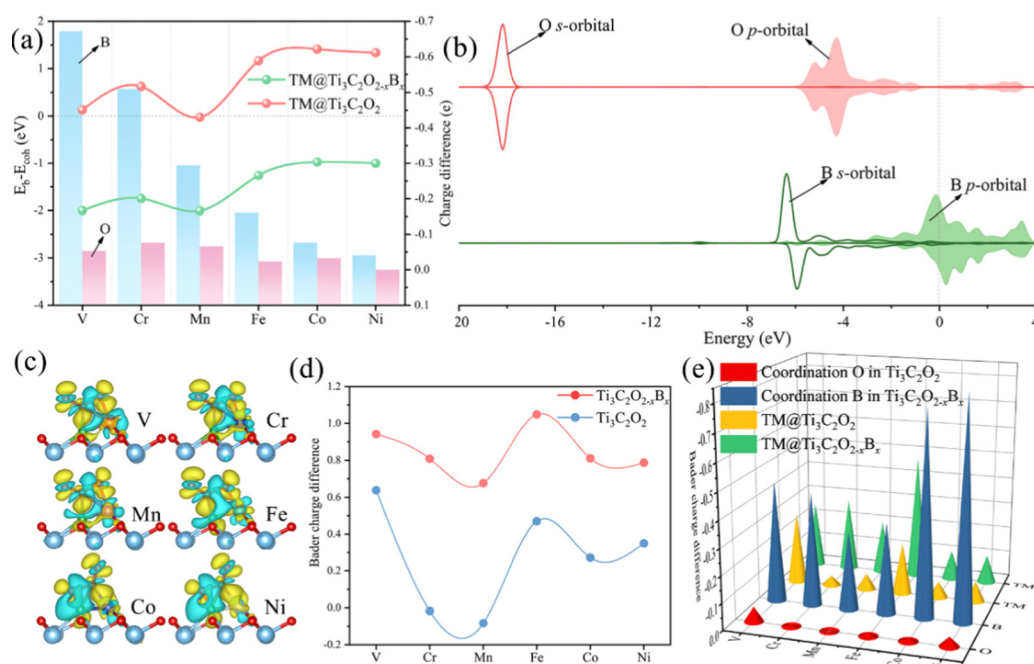


Fig. 2 (a) The $E_b - E_{\text{coh}}$ values of $\text{TM}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{TM}@\text{Ti}_3\text{C}_2\text{T}_x$ (line chart), charge differences of B and O atoms before and after anchoring a single TM atom (histogram). (b) PDOS (projected density of states) of the O or B surface atoms before fixing the TM single atoms. (c) Variation of charge density differences of $\text{TM}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ with a CO_2 adsorption configuration. The yellow region denotes charge accumulation and the blue region indicates charge depletion, where the isosurface value is set to be $0.03 \text{ e } \text{\AA}^{-3}$. (d) Bader charge differences of the C atom in CO_2 before and after adsorption on $\text{TM}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{TM}@\text{Ti}_3\text{C}_2\text{T}_x$. (e) Bader charge differences of TM atoms, the coordination O atom in $\text{TM}@\text{Ti}_3\text{C}_2\text{O}_2$ and the coordination B atom in $\text{TM}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ before and after CO_2 adsorption.

and $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$. Previous studies, both theoretical and experimental, have identified that the H1 site is a stable adsorption site for metal atoms (Fig. 1c and d).^{51,55}

Thermodynamically, the aggregation of the dispersed TM atoms depends on the bond strength between the TM and the substrate as well as the recombination ability of metal atoms. Therefore, the stability of SACs can be measured using the values of $E_b - E_{\text{coh}}$ of the TM and substrate (eqn (1) and (2)): the smaller this value the more thermodynamically stable the structure of SACs (Table S1, ESI†). For all of the TM atoms, the value of $E_b - E_{\text{coh}}$ for $\text{TM}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ is lower than that for

$\text{TM}@\text{Ti}_3\text{C}_2\text{O}_2$, suggesting that the presence of B is advantageous for stabilizing TM atoms (Fig. 2a). This can be explained by the fact that B atoms bind TM atoms through substantial partial charge transferred from the latter to the former. When the B atom is doped into the $\text{Ti}_3\text{C}_2\text{O}_2$ substrate, two electrons are transferred from the latter to the former, leaving the partial 2p orbital with lack of electrons. The empty 2p orbital helps the B atom gaining more electrons from the TM single atom than from O atom (Fig. 2a), leading to a thermodynamically more stable structure.

As shown in Fig. 2b, since the localized 2s orbital of the O atom is in the deep energy level, it is difficult for this orbital to

participate in bonding. The 2p orbital of the O atom is almost below the Fermi level, implying that it is fully occupied. The O atom with a saturated electronic structure is not conducive for receiving electrons from TM atoms. Compared with the O atom, the 2s orbital of the B atom is localized in a shallower energy level that is adjacent to its 2p orbital. The 2s orbital of the B atom and the d orbitals of the TM atom have obvious overlaps in the energy level ranging from -6 to 0 eV, indicating intense interactions between the TM and B atoms (Fig. S1, ESI†).

A part of the 2s orbital of the B atom is at the delocalization state and slightly contributes to the electronic state at the Fermi level, demonstrating the hybridization of B s-p orbitals. The partial 2p orbital of the B atom is above the Fermi level and the unoccupied 2p orbital can accept electrons from TM atoms to form strong-polarized covalent TM-B bonds. In addition, we have calculated the area of the H1 site, which is defined as the triangular area enclosed by three oxygen atoms in $\text{Ti}_3\text{C}_2\text{O}_2$ or two oxygen atoms and one boron atom in $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (Fig. S2, ESI†). The calculated area of the H1 site of $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (4.21 \AA^2) is larger than that of $\text{Ti}_3\text{C}_2\text{O}_2$ (4.11 \AA^2). The larger area in the former provides more spacious H1 sites so that TM atoms can be effectively embedded into the flat substrate of $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$. Since our results suggest that $\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ has an outstanding ability to fix TM single atoms, all the six kinds of metal are considered in the following CO_2 RR investigation.

3.2. CO_2 electroreduction reaction on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ SACs

3.2.1. Initial activation of CO_2 . One method of effective activation of CO_2 is changing its linear structure ($\text{O}=\text{C}=\text{O}$) to a V-shaped ($\angle \text{OCO} < 180^\circ$) structure, where partially negatively charged CO_2 is generated due to the injection of electrons into its antibonding $2\pi_u^*$ orbital. As shown in Fig. S3 (ESI†), CO_2 binds to $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (TM is V, Cr, Mn or Fe) through a B-C-TM-O configuration, but to $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (TM is Co or Ni) through an O-B-C-TM configuration. The linear $\text{O}=\text{C}=\text{O}$ is bent substantially and its C atom binds to the TM atom and the nearby B atom in both configurations. These bonding configurations indicate that the B atom can participate in the CO_2 adsorption. In the case of $\text{TM@Ti}_3\text{C}_2\text{O}_2$ (Fig. S3, ESI†), CO_2 binds to TM (TM is V, Fe, Co or Ni) through a bidentate C-TM-O configuration with a bent structure, but binds to TM (TM is Cr or Mn) through a monodentate O-M configuration

with a linear structure. This reveals that $\text{TM@Ti}_3\text{C}_2\text{O}_2$ (TM is V, Fe, Co or Ni) also has the ability to activate CO_2 molecules. However, for all the $\text{TM@Ti}_3\text{C}_2\text{O}_2$ considered, CO_2 molecules only bind to the TM atom, implying that the O atom of the functional groups does not participate in the CO_2 activation directly. As can be seen in Table 1, the bending angle of the CO_2 adsorbed on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ is smaller than that on $\text{TM@Ti}_3\text{C}_2\text{O}_2$, suggesting a more significant CO_2 activation in the former.

These phenomena we discussed above indicate that the doping of B facilitates the activation of CO_2 through participating in its adsorption process. To understand the emerging role of the B atom in CO_2 activation, we examine the electronic structures of $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{TM@Ti}_3\text{C}_2\text{O}_2$ through a Bader charge analysis. The images of the charge density difference of the CO_2 adsorbed $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ are shown in Fig. 2c. As seen, charge accumulated in the CO_2 $2\pi_u^*$ antibonding orbital around the C atom, leading to the activation of CO_2 molecules. In contrast, charge depleted around B and TM atoms, confirming that the electrons involved in the activation of CO_2 come from these atoms. To give more details of charge transfer density, Bader effective charge is investigated. When adsorbed on the surface, the C atom of CO_2 in the case of $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ can gain more electrons than that in the case of $\text{TM@Ti}_3\text{C}_2\text{O}_2$ (Fig. 2d).

Generally, the strong electronegativity of oxygen shifts the focus of electronic clouds on both ends of CO_2 molecules, yielding an electron-deficient carbon center. An injection of electron into the C atom is an effective way to activate CO_2 molecules. The observation that $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ has a more activated CO_2 is consistent with the results of calculations. Bader effective charge analysis (Fig. 2e) also shows that B and TM atoms can provide enough electrons for injection into CO_2 molecules. In particular, in the cases of $\text{Co@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{Ni@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$, most of the injected electrons come from the B atom. When CO_2 molecules are activated on $\text{TM@Ti}_3\text{C}_2\text{O}_2$, the $\text{Ti}_3\text{C}_2\text{O}_2$ substrate can provide extra electrons to activate CO_2 , which is consistent with our previous result.⁵⁶ But an indirect participation mode provides less electron than a direct participation.

The reaction activity of the doped B atom can be ascribed to its occupied 2p orbitals with s-p hybridization. Because the electronegativity of B is weaker than that of C, the electrons in the occupied orbitals of the B atom can transfer easily to the C atom of the CO_2 molecule. In contrast, the full 2p-orbital occupation electronic structure of the O atom leads to a strong chemical stability. The stronger electronegativity of the O atom also makes it difficult for the C atom to grab its electrons. To gain insights into the chemisorption of CO_2 on the active site of $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$, the bonding states of CO_2 on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ were investigated. According to the projected density of states (PDOS) for an isolated CO_2 molecule (Fig. 3a-d and Fig. S4, ESI†), the empty orbital above the Fermi level is the antibonding $2\pi_u^*$ orbital.

According to the PDOS of bare $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$, TM 3d overlaps with B 2p, indicating the bonding of TM and B atoms.

Table 1 Binding energy (BE), binding distance (BD), and bond angle (BA) of CO_2 adsorbed on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{TM@Ti}_3\text{C}_2\text{O}_2$ in the most stable configuration

Metal	$\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$				$\text{Ti}_3\text{C}_2\text{O}_2$		
	BE (eV)	BD (\AA)		BA ($\angle \text{OCO}$, $^\circ$)	BE (eV)	BD (\AA)	BA ($\angle \text{OCO}$, $^\circ$)
		TM-C	B-C			TM-C	
V	-1.49	2.12	1.72	128.7	-0.62	2.01	140.7
Cr	-1.03	2.10	1.74	130.5	-0.65	—	179.7
Mn	-0.52	2.15	1.81	136.5	-0.48	—	179.4
Fe	-0.73	2.48	1.62	126.5	-0.27	2.06	151.9
Co	-1.34	1.86	1.82	125.3	-0.09	2.09	158.9
Ni	-1.62	1.89	1.82	126.8	-0.25	1.98	153.5

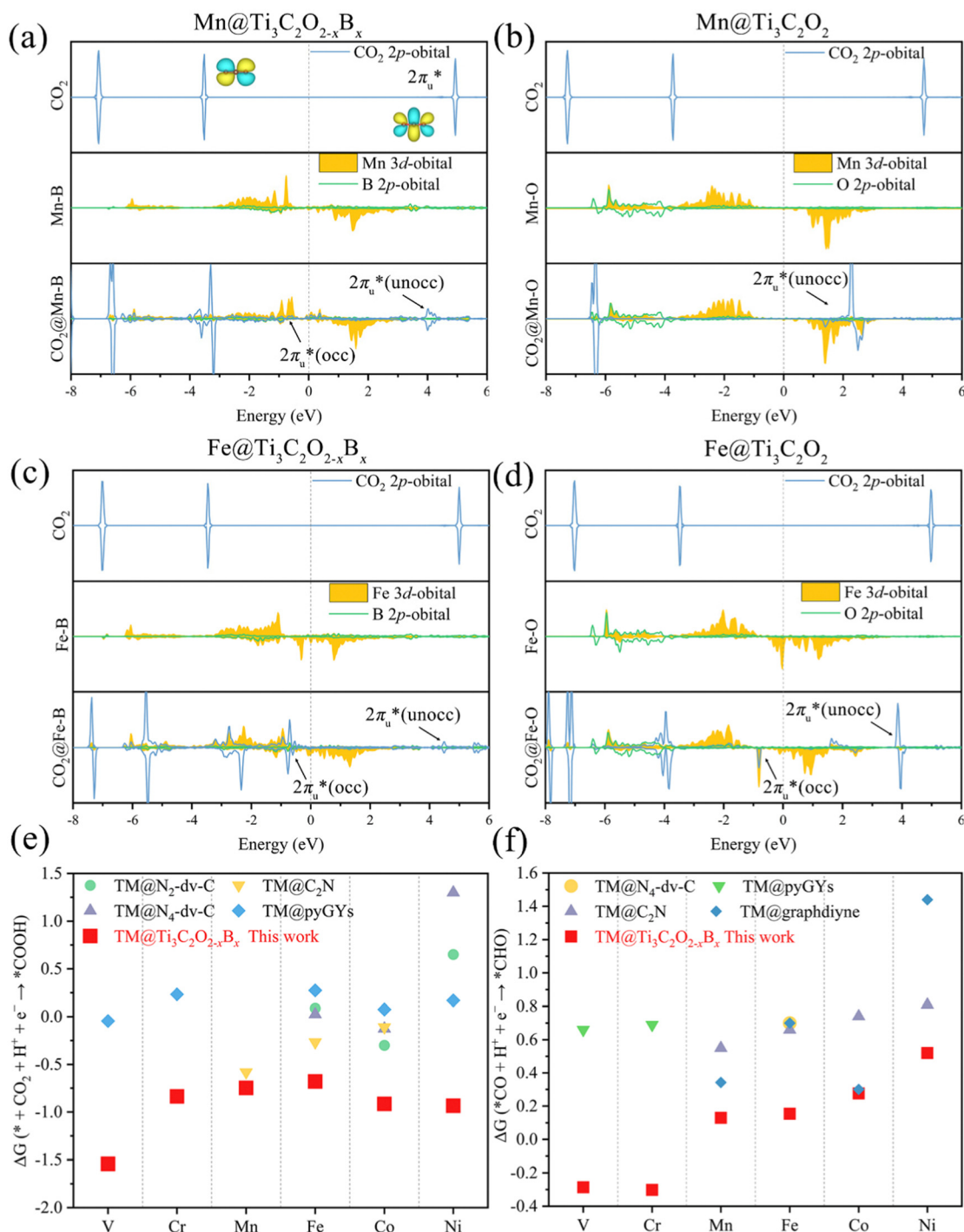


Fig. 3 The PDOS of CO₂ 2p, TM 3d, and coordination atoms (the O atom in TM@Ti₃C₂O₂ and the B atom in TM@Ti₃C₂O_{2-x}B_x) of the 2p orbital before and after CO₂ adsorption on (a) Mn@Ti₃C₂O_{2-x}B_x, (b) Mn@Ti₃C₂O₂, (c) Fe@Ti₃C₂O_{2-x}B_x and (d) Fe@Ti₃C₂O₂. A comparison of the Gibbs free energy changes of (e) the 1st (CO₂ + H⁺ + e⁻ → *COOH) and (f) the 3rd (CO + H⁺ + e⁻ → *CHO) hydrogenation reactions for TM@Ti₃C₂O_{2-x}B_x and other SACs (TM@N₂-dv-C,⁷⁴ TM@N₄-dv-C,⁷⁴ TM@C₂N,⁷⁵ TM@pyGYs,⁷⁶ and TM@graphdiyne⁷⁷) for the CO₂ reduction.

After its adsorption on TM@Ti₃C₂O_{2-x}B_x, the antibonding 2π_u^{*} of CO₂ splits into two parts: one above and one below the Fermi level. The states above the Fermi level are contributed by the unoccupied TM 3d-CO₂ 2π_u^{*} and B 2p-CO₂ 2π_u^{*} hybridization orbitals, while the overlapping of occupied TM 3d-CO₂ 2π_u^{*} and B 2p-CO₂ 2π_u^{*} hybridization orbitals can be found at

-2 to 0 eV, just below the Fermi level. After adsorbing on TM@Ti₃C₂O₂, the antibonding 2π_u^{*} orbital of CO₂ also splits into one part above the Fermi level and one part below it. Notably, the peak intensity of the unoccupied CO₂ 2π_u^{*} orbital when CO₂ adsorbed on TM@Ti₃C₂O_{2-x}B_x decreases more sharply than that when CO₂ adsorbed on TM@Ti₃C₂O₂. This implies that

TM@Ti₃C₂O_{2-x}B_x provides more electrons that can be injected into the 2π_u^{*} orbital, which is consistent with the results of Bader charge analysis.

Besides CO₂ activation ability, another key issue in CO₂RR catalyst design is to obtain an energy-rich product (CH₄), which is closely related to the adsorption Gibbs free energy of radical reaction intermediates and that of the elementary CO₂ reduction reaction. CO₂RR can be differentiated by the types of intermediates formed in the initial protonation steps of CO₂: formate (*OCHO) and carboxyl (*COOH). The calculated Gibbs free energy changes (Δ*G*) for the formation of *COOH and *OCHO (*CO₂ + H⁺ + e⁻ → *COOH or *OCHO) on TM@Ti₃C₂O_{2-x}B_x are summarized in Fig. S5 (ESI†). This figure shows that the 1st hydrogenation (H⁺/e⁻) reactions are energetically prone to occur on the O atom rather than on the C atom, suggesting that the formation of *COOH is more energetically preferred than that of *OCHO. The calculated Δ*G* shown in Fig. 3e reveals that the Δ*G* of the first hydrogenation reaction to form *COOH on TM@Ti₃C₂O_{2-x}B_x is smaller than that in previous studies, illustrating the superior CO₂ activation ability of TM@Ti₃C₂O_{2-x}B_x.

3.2.2. Origin of CO activity on TM@Ti₃C₂O_{2-x}B_x. According to previous studies, *COOH will always lead to the formation of *CO *via* releasing one H₂O molecule, and further reduce to intermediate *CHO.^{14,74,75} Besides *COOH, *CO is another crucial intermediate in the CO₂RR to form CH₄, the adsorption energy of which is closely related to CO₂RR performance. The weak CO adsorption on SACs results in direct desorption of CO under low electrode potentials. Therefore, a stably adsorbed CO on SACs may improve its further reduction towards CH₄.

The relatively higher CO adsorption strength on TM@Ti₃C₂O_{2-x}B_x (Table S2, ESI†) results in poor activity for CO direct desorption under low electrode potentials. Due to the outstanding stability of CO, the hydrogenation reaction needs too much energy to reduce *CO. If the Δ*G* of the reaction *CO + H⁺ + e⁻ → *CHO is larger than that of the reaction *CO → CO + *, CO is inclined to desorb rather than further hydrogenation. However, a high voltage is still needed to remove *CO from the surface if the adsorption of CO on the catalyst is strong. In general, the lower hydrogenation Δ*G* of the reaction *CO + H⁺ + e⁻ → *CHO is necessary for the further CO reduction process. Therefore, the activation of the *CO intermediate is a key criterion to evaluate the performance of the CO₂RR on SACs.

To verify the performance of TM@Ti₃C₂O_{2-x}B_x in *CO reduction, we calculated the Δ*G* of the reaction *CO + H⁺ + e⁻ → *CHO, along with the Δ*G* of other SACs. Having examined the available results for the CO₂RR of SACs, we can find that *CO + H⁺ + e⁻ → *CHO is the potential-determining step in most of these studies. As shown in Fig. 3f, the calculated Δ*G* for the reaction *CO + H⁺ + e⁻ → *CHO on TM@Ti₃C₂O_{2-x}B_x is smaller than the corresponding values of Δ*G* reported in previous studies, which verifies the superior CO activation ability of TM@Ti₃C₂O_{2-x}B_x. According to the adsorption configuration of *CO on TM@Ti₃C₂O_{2-x}B_x, the B atom also participates in the adsorption process (Fig. 4a), demonstrating its important role in CO activation. To gain insights into the

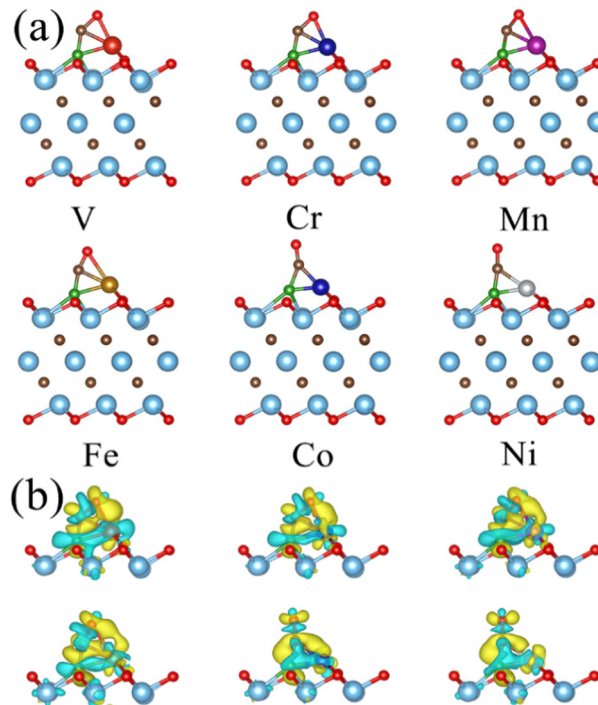


Fig. 4 (a) Optimized geometric structures of CO adsorbed on TM@Ti₃C₂O_{2-x}B_x. (b) Variation in the difference of charge density of TM@Ti₃C₂O_{2-x}B_x with a CO adsorption configuration. The yellow region denotes charge accumulation and the blue region indicates charge depletion, where the isosurface value is set to be 0.03 e Å⁻³.

combined effect of multi-active site TM and B atoms in ameliorating the CO reduction, the orbital and charge transfer of TM and B atoms during the adsorption of *CO is also investigated. The configuration of the *CO adsorbed on TM@Ti₃C₂O₂ is used for comparison.

Accepting the lone-pair electrons from the σ_g orbital of CO and donating electrons to the antibonding π_g^{*} orbitals of CO are the main ways to weaken the C≡O triple bond, known as the “acceptance–donation” process.^{78,79} Fig. 4b and Fig. S6 (ESI†) illustrate the calculated charge density differences for the CO adsorption configuration on TM@Ti₃C₂O_{2-x}B_x and TM@Ti₃C₂O₂. Charge accumulation and depletion can be observed for both CO molecules and B–TM atoms on TM@Ti₃C₂O_{2-x}B_x. This phenomenon is actually in perfect accordance with the “acceptance–donation” process mentioned above. A comparison between the area of the isosurface of CO adsorbed on TM@Ti₃C₂O_{2-x}B_x and that on TM@Ti₃C₂O₂ reveals that the charge transfer in the former is more obvious than that in the latter.

To further elucidate the bonding nature, the PDOS of the adsorption of CO on TM@Ti₃C₂O_{2-x}B_x and that on TM@Ti₃C₂O₂ are investigated (Fig. 5 and Fig. S7, ESI†). After adsorbing on TM@Ti₃C₂O_{2-x}B_x and TM@Ti₃C₂O₂, the σ_g orbital of CO in a deep level obviously splits and moves to deeper energy states, indicating the interaction between the CO σ_g orbital and SACs. It can be seen that a part of the π_g^{*} orbitals of CO unoccupied (unocc) shifts upward to the Fermi level and the



Fig. 5 The PDOS of CO 2p, TM 3d, and coordination atoms (the O atom in $\text{TM@Ti}_3\text{C}_2\text{O}_2$ and the B atom in $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$) of the 2p orbital before and after CO_2 adsorption on (a) $\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$, (b) $\text{Mn@Ti}_3\text{C}_2\text{O}_2$, (c) $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and (d) $\text{Fe@Ti}_3\text{C}_2\text{O}_2$.



Fig. 6 (a) Bader charge differences of the C atom in CO before and after adsorption on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{TM@Ti}_3\text{C}_2\text{T}_x$. (b) Bader charge differences of TM atoms and the B atom on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ before and after CO adsorption.

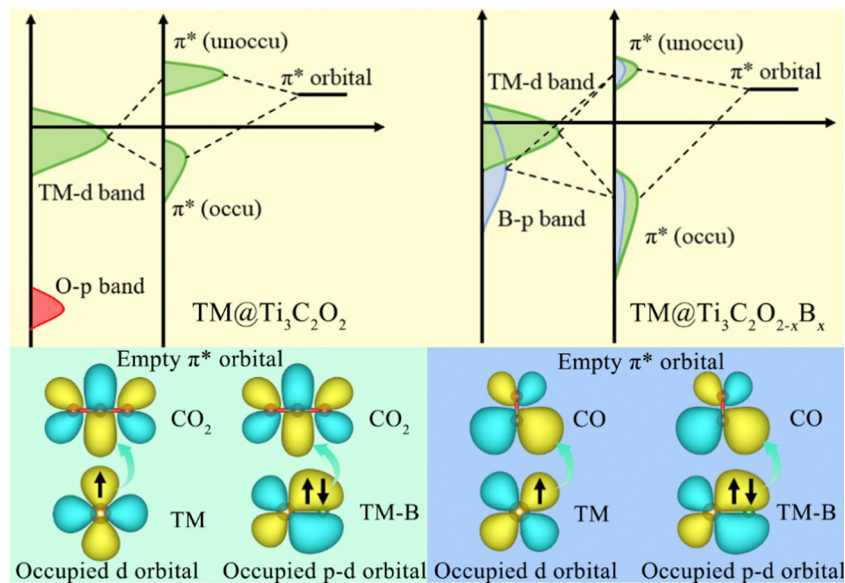


Fig. 7 Schematic diagrams of CO₂ molecule and CO intermediate activation mechanisms on TM@Ti₃C₂O₂ and TM@Ti₃C₂O_{2-x}B_x.

other part moves downward to the Fermi level near -2 eV to form a π_g^* occupied (occ) orbital. Compared with that of TM@Ti₃C₂O₂, the peak intensity of the unoccupied CO π_g^* antibonding orbital decreases more sharply when CO is adsorbed on TM@Ti₃C₂O_{2-x}B_x, indicating the better CO activation performance of TM@Ti₃C₂O_{2-x}B_x. The Bader charge analysis shows that the C atom from CO can gain more electrons from TM@Ti₃C₂O_{2-x}B_x (Fig. 6a and b), which is consistent with the above PDOS analysis.

Owing to its weak electronegativity and not fully occupied orbital structure, the electrons in the occupied orbitals of the B atom can easily transfer to the empty orbitals of the CO₂ molecule and CO intermediate. Based on the results gathered, it can be reasonably inferred that a combined effect between the single TM atom and B atom is present on the CO₂RR (Fig. 7). The TM-to-C and B-to-C π -back bonding contribute to the activation of the CO₂ molecules and CO intermediate. Enough electrons from the single TM atom and B atom can be injected into CO₂ and CO molecules through a direct bonding interaction. Therefore, the B-doping strategy efficiently controls the charge-orbital state of the coordination atom on the surface of Ti₃C₂O_{2-x}MXene, alleviating remarkably the difficulty of activating CO₂ and CO molecules.

3.2.3. Subsequent protonation steps toward CO₂RR production. The excellent activation performance endows TM@Ti₃C₂O_{2-x}B_x SACs with great potential for the CO₂RR. The outstanding CO activation performance can also help its further reduction reaction towards methane on TM@Ti₃C₂O_{2-x}B_x SACs. Using the electronic potential energy surface (PES) as a baseline, we examine the subsequent protonation steps *via* the *COOH intermediate to form final product CH₄(g) in this section.

Fig. 8 shows the variation of the calculated free energy with the most favorable pathway for the reduction of CO₂

on TM@Ti₃C₂O_{2-x}B_x SACs and the corresponding structures of the reaction intermediates are presented in Fig. S8 (ESI[†]). Owing to the excellent CO₂ activation performance of TM@Ti₃C₂O_{2-x}B_x, the hydrogenation of the adsorbed CO₂ to form *COOH is exothermic with a negative ΔG . The 3rd hydrogenation (H^+/e^-) reaction of *CO occurs, generating *CHO. *CO is more likely to undergo further reduction on all the TM@Ti₃C₂O_{2-x}B_x because the ΔG of the protonation step ($*CO + H^+ + e^- \rightarrow *CHO$) is more negative than that of CO desorption. Therefore, we concluded that the production of CO can be effectively prohibited in the CO₂ reduction. The small ΔG of the 3rd hydrogenation also implies that the adsorbed *CO, which is generally regarded as the poison on metallic electrodes, is less likely to poison the active sites in our case.

For the subsequent elementary reaction steps passing through the CHO* intermediate to generate CH₄, it is worth mentioning that the desorption of CH₂O from TM@Ti₃C₂O_{2-x}B_x is significantly endothermic due to their large negative adsorption energies (Table S2, ESI[†]), indicating that the production of methanol (CH₂O) can be effectively prevented during the reduction of CO₂ catalyzed by TM@Ti₃C₂O_{2-x}B_x. In addition, *CH₃O is more likely to be further reduced to *O, releasing one CH₄ molecule on TM@Ti₃C₂O_{2-x}B_x (TM is V, Cr or Mn). Therefore, the production of CH₃OH can be avoided effectively. For the case of TM@Ti₃C₂O_{2-x}B_x (TM is Fe, Co or Ni), the desorption ΔG of CH₃OH is appreciably large so that the desorption of CH₃OH is difficult. After *CH₃OH is formed on TM@Ti₃C₂O_{2-x}B_x, its C atom will be attacked by a proton to form CH₄(g) and *OH. This suggests that CH₄ is the primary reduction product on all the TM@Ti₃C₂O_{2-x}B_x examined, indicating its great selectivity for catalyzing the CO₂ reduction to form CH₄ with excellent abilities to limit the production of CO(g), CH₂O(g), and CH₃OH(l).



Fig. 8 Variation of the Gibbs free energy for the reduction reaction of CO_2 on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (TM is V, Cr, Mn, Fe, Co or Ni) at 0 V with RHE.

To assess the durability of $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ SACs in practical application, we have further investigated the protonation of remaining $^*\text{OH}$ as well as the release of the adsorbed H_2O molecule (8th hydrogenation reaction). The results obtained show that the $^*\text{OH}$ on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (TM is V or Cr) is harder to be removed; the uphill values of ΔG are 1.47 and 0.96 eV, respectively. The positive ΔG indicates that $^*\text{OH}$ is more likely to occupy active sites by forming TM-OH , which hinders the next CO_2 reduction process. For other $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (TM is Mn, Fe, Co or Ni), the formation of $^*\text{CHO}$ ($^*\text{CO} + \text{H}^+ + \text{e}^- = ^*\text{CHO}$) and $^*\text{CH}_3\text{O}$ ($^*\text{CH}_2\text{O} + \text{H}^+ + \text{e}^- = ^*\text{CH}_3\text{O}$) is both endothermic and the rest of the elementary steps are exothermic.

Except $\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$, the uphill value of ΔG for the formation of $^*\text{CH}_3\text{O}$ on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (TM is Fe, Co or Ni) is larger than that for the formation of $^*\text{CHO}$. Therefore, the formation of $^*\text{CHO}$ and $^*\text{CH}_3\text{O}$ is the potential-determining step (PDS) for $\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (TM is Fe, Co or Ni), respectively. The PDS and the values of U_L predicted by DFT are summarized in Table 2. Based on the U_L calculated from the primary PDS, it can be inferred that $\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ have a higher activity for CH_4 production, where their limiting potentials are 0.13 and 0.3 V, respectively, lower than those of other $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$.

Table 2 DFT-predicted potential determining steps (PDSs), U_L values, and possible products on $\text{TM@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$

Catalyst	Potential determining step	U_L (V)	Production
$\text{V@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$	$^*\text{OH} + \text{H}^+ + \text{e}^- \rightarrow \text{H}_2\text{O}(\text{l})$	1.47	CH_4
$\text{Cr@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$	$^*\text{OH} + \text{H}^+ + \text{e}^- \rightarrow \text{H}_2\text{O}(\text{l})$	0.96	CH_4
$\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$	$^*\text{CO} + \text{H}^+ + \text{e}^- \rightarrow ^*\text{CHO}$	0.13	CH_4
$\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$	$^*\text{CH}_2\text{O} + \text{H}^+ + \text{e}^- \rightarrow ^*\text{CH}_3\text{O}$	0.30	CH_4
$\text{Co@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$	$^*\text{CH}_2\text{O} + \text{H}^+ + \text{e}^- \rightarrow ^*\text{CH}_3\text{O}$	0.55	CH_4
$\text{Ni@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$	$^*\text{CH}_2\text{O} + \text{H}^+ + \text{e}^- \rightarrow ^*\text{CH}_3\text{O}$	0.68	CH_4

We conclude that $\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ and $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ have a remarkable selectivity and activity for the production of $\text{CH}_4(\text{g})$. Considering the effect of the presence of water under real reaction conditions, we calculated the free energy diagram of the initial screening for $\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}$ and $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}$ in the implicit solvent model using VASPsol (Fig. 9a and b). Notably, the levels of U_L for $\text{Mn@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}$ and $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}$ increased to 0.51 and 0.40 V, respectively. The overpotential of $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}$ (0.40 V) suggests that it can be a promising CO_2RR catalyst for producing CH_4 . This level of limiting potentials is comparable or lower than those of the widely employed transition metal-based CO_2RR electrocatalysts^{8,14,80} and other bimetallic electrocatalysts work on CO_2 conversion to CH_4 (Table S3, ESI[†]).

3.3. Stability of $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ in solution and competition between the CO_2RR and the HER

Apart from CO_2 adsorption, the competitive adsorption of CO_2RR and H on each active center is also a significant factor affecting CO_2RR performance. The free energy of hydrogen adsorption (ΔG_{H}) provides a measure for evaluating HER activity. A material is an optimal HER electrocatalyst if its ΔG_{H} value is close to zero. However, a highly negative ΔG_{H} indicates the excessively stable adsorption of H on the active sites, thereby preventing the CO_2RR . Therefore, active sites with a large ΔG_{H} are desirable for CO_2RR electrocatalysts. Assuming a Volmer and Tafel mechanism, we calculated the ΔG for the formation of $^*\text{COOH}$ ($^* + \text{CO}_2 + \text{H}^+ + \text{e}^- \rightarrow ^*\text{COOH}$) and that for the formation of $^*\text{H}$ ($\text{H}^+ + \text{e}^- \rightarrow ^*\text{H}$) in the HER. According to Brønsted–Evans–Polanyi (BEP) relations, a reaction with a more negative ΔG is expected to encounter a smaller reaction barrier and, therefore, is more favored to occur.

The adsorption sites for the H atom mainly locate on the bridge site between the TM and B atoms on $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ (Fig. 9c). The calculated ΔG_{H} of $\text{Fe@Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ is -0.48 eV,



Fig. 9 Free energy diagrams of the CO₂RR on (a) Mn@Ti₃C₂O_{2-x}B and (b) Fe@Ti₃C₂O_{2-x}B after taking into account of solvation effects. (c) Variation of the Gibbs free energy changes (ΔG) for the first protonation step of the CO₂RR with HER. Inset: the optimized geometric structures of H adsorbed on Fe@Ti₃C₂O_{2-x}B. (d) Variations of energy and temperature with the AIMD simulation time for Fe@Ti₃C₂O_{2-x}B, where the AIMD simulation lasts for 10 ps at 500 K. (e) Variations of energy and temperature with the AIMD simulation time for Fe@Ti₃C₂O_{2-x}B, an explicit solvent model is used to simulate the acidic aqueous solution. The AIMD simulation lasts for 10 ps at 330 K. (f) Calculated Pourbaix diagram of Fe@Ti₃C₂O_{2-x}B. Gray dashed lines denote the water redox potentials.

which is more positive than the ΔG_{*COOH} of the CO₂RR (Fig. 9c). This suggests that Fe@Ti₃C₂O_{2-x}B_x contains inactive sites for the HER with a $|\Delta G_{H}|$ of around 0.5 eV and exhibits a higher selectivity toward the CO₂RR than the HER.

To assess the durability of MXene catalysts in practical applications, we investigated the stability of Fe@Ti₃C₂O_{2-x}B_x via AIMD simulations. As seen in Fig. 9d, both the total energy and the temperature oscillate near the initial conditions and the geometric structures of Fe@Ti₃C₂O_{2-x}B_x are preserved well

within 10 ps. An explicit solvent model is also applied to simulate the solution environment, and AIMD simulations are performed to study the thermal stability of Fe@Ti₃C₂O_{2-x}B_x in a solution at 330 K for 10 ns.

The results obtained reveal that the structure of Fe@Ti₃C₂O_{2-x}B_x can remain stable in an aqueous solution (Fig. 9e). Therefore, it can be inferred that Fe@Ti₃C₂O_{2-x}B_x can serve as an efficient CO₂RR electrocatalyst under ambient conditions with a high stability. Further analysis of the Pourbaix diagram



Fig. 10 The limiting potential as a function of the valence state of TM atoms fixed on (a) Ti₃C₂O_{2-2x}B_x and (b) Ti₃C₂O₂, Ti₃C₂O_{2-2x}N_x.

of Fe@Ti₃C₂O_{2-2x}B_x (Fig. 9f) has also been carried out. It is a useful tool to identify the stable state of catalysts in water as a function of pH and applied potential. The oxidation potential of Fe@Ti₃C₂O_{2-2x}B_x at pH 0 (black line) is less negative than the U_L value of Fe@Ti₃C₂O_{2-2x}B_x (red line), indicating that Fe@Ti₃C₂O_{2-2x}B_x could possess superior electrochemical stability against surface oxidation under working conditions.

3.4. An activity descriptor for CO₂ reduction on Ti₃C₂T_x-based SACs

It is beneficial to search an activity descriptor to establish a “volcano” relationship on catalytic behaviors so that we can further roughly estimate the effect of other potential dopants on the CO₂RR performance. It is generally agreed that the electronic structures of catalytic centers would greatly influence the electron transfer and reaction energy in catalytic processes. Pan and coworkers have found that the valence state of TMs could be used as an activity descriptor, wherefore the limiting potential towards different products as a function of the valence state of TMs⁸¹ because the TM atoms can provide both empty d orbitals to accept electrons to form a coordination bond and lone electrons to form a covalent bond.⁸² Hence, the valence state of TMs would directly influence the bonding strength with the intermediates.

The valence state of different TM atoms is defined as the number of electrons transferred from the TM atom to the substrate. Using the valence state of TMs as the descriptor, we find that a volcano trend is observed for TM@Ti₃C₂O_{2-2x}B_x (Fig. 10a), meaning this descriptor is possible to estimate the catalytic performance of the TM@Ti₃C₂O₂ system. Therefore, a brief calculation of N-doped Ti₃C₂O₂-based SACs (TM@Ti₃C₂O_{2-2x}N_x) is carried out to verify the accuracy of this descriptor. The adsorption configuration of CO₂ and CO on TM@Ti₃C₂O_{2-2x}N_x (Fig. S9, ESI[†]) indicates that the nitrogen atom is only used as the coordination atom to control the electronic structure of the TM. Generally, the potential-limiting step of CO₂ reduction is the hydrogenation of *COOH to *CO or the hydrogenation of *CO to *CHO.^{83,84} As for the TM@Ti₃C₂O₂ system studied in our work, *OH + H⁺ + e⁻ → H₂O(l) is the

potential-limiting step on V@Ti₃C₂O₂ and Cr@Ti₃C₂O₂. Hence, we have calculated these key reaction steps on TM@Ti₃C₂O_{2-2x}N_x and chosen the step with the largest reaction energy barrier as the potential-limiting step.

The result shows that a volcano trend is observed with the optimal scope of the valence state ranging from ~0.8 e to ~1.1 e (Fig. 10b). We can see that the optimal scope of the valence state is different between TM@Ti₃C₂O_{2-2x}B_x and TM@Ti₃C₂O₂ (TM@Ti₃C₂O_{2-2x}B_x). Because the boron atom can participate in CO₂ reduction, the scaling relations between the adsorption strength of reaction intermediates have been broken. This simple descriptor well explains the catalytic performance of the TM@Ti₃C₂O₂ and doped TM@Ti₃C₂O₂ SACs and suggests that a certain moderate degree of electron transfer from TM atoms to the substrate will be expected to promote the overall catalytic activity. Before using this descriptor, we need to determine whether the coordination atoms participate in the reaction to ensure that the same linear relationship exists. Moreover, this descriptor can only roughly estimate the CO₂RR performance of catalysts. A series of subsequent tests, such as the competition test, stability test and solvation effect should be considered to further screening out the most suitable catalysts.

4. Conclusions

In summary, by applying well-defined first-principles calculations, we have investigated the single-atom catalysts of Ti₃C₂O_{2-2x}B_x monolayers for CO₂ electrochemical reduction reactions. The results obtained suggest that the B atom not only serves as the coordination atom to anchor single TM atoms but also plays an important role in the CO₂ electrochemical reduction process. The strategy of doping B regulates the charge-orbit distribution of the surface functional groups of Ti₃C₂O₂. By analyzing charge transfer and DOS calculation results, the combined effect of the single TM atom and the adjacent B atom in Ti₃C₂O_{2-2x}B_x on the CO₂RR has been

demystified. The empty orbitals in B atom hybridization help it in gaining more electrons from single TM atoms, forming strong-polarized covalent TM–B bonds, which is beneficial for anchoring single TM atoms on the $\text{Ti}_3\text{C}_2\text{O}_2$ surface. The electrons provided by the occupied orbitals in B atom hybridization are effectively transferred to the π -antibonding orbitals of CO_2 and CO, thereby helping single TM atoms to activate the CO_2 molecules and *CO intermediate. Especially, the activation of the *CO intermediate induced by the combined effect effectively lowers the reaction energy barrier of the protonation step of *CO, facilitating further CO reduction towards CH_4 .

Our results also demonstrated that the activated CO_2 can be further reduced to CH_4 by $\text{TM}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$, in which $\text{Fe}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ exhibits a better catalytic efficiency with a low limiting potential of ~ 0.40 V. In addition, AIMD calculations reveal that $\text{Fe}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ has a high stability. Inspired by the successful preparation of $\text{Ru}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ SACs,⁵⁵ the practical application of $\text{Fe}@\text{Ti}_3\text{C}_2\text{O}_{2-x}\text{B}_x$ in CO_2 capture and reduction can be reasonably anticipated in the near future. We conclude that our work provides atomic-level insights into the combined effect between the embedded single metal atom and the innate atom in a substrate and provides an important strategy for the design of MXenes as novel CO_2RR catalysts.

Conflicts of interest

There are no conflicts to declare.

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