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Zirconia-based nanomaterials: recent developments in synthesis and applications

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In the last decade, the whole scientific community has witnessed great advances and progress in the various fields of nanoscience. Among the different nanomaterials, zirconia nanomaterials have found numerous applications as nanocatalysts, nanosensors, adsorbents, etc. Additionally, their exceptional biomedical applications in dentistry and drug delivery, and interesting biological properties, viz. anti-microbial, antioxidant, and anti-cancer activity, have further motivated the researchers to explore their physico-chemical properties using different synthetic pathways. With such an interest in zirconia-based nanomaterials, the present review focuses systematically on different synthesis approaches and their impact on the structure, size, shape, and morphology of these nanomaterials. Broadly, there are two approaches, viz., chemical synthesis which includes hydrothermal, solvothermal, sol-gel, microwave, solution combustion, and co-precipitation methods, and a greener approach which employs bacteria, fungus, and plant parts for the preparation of zirconia nanoparticles. In this review article, the aforementioned methods have been critically analyzed for obtaining specific phases and shapes. The review also incorporates a detailed survey of the applications of zirconia-based nanomaterials. Furthermore, the influence of specific phases, morphology, and the comparison with their counterpart composites for different applications have also been included. Finally, the concluding remarks, prospects and possible scope are given in the last section.

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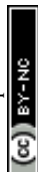


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1 Introduction

Zirconium is a tough transition metal which has found boundless applications when present in its oxide form, known as zirconia or zirconium oxide. Zirconia is also known as ceramic steel owing to its intrinsic hardness, abrasiveness, high

melting point, and low frictional resistance. In the nanoscale, it becomes immensely valuable owing to its high thermal stability, luminescence, refractive index, chemical stability, high specific area, biocompatibility, and exhibiting significant antibacterial, antioxidant, and antifungal properties. Such outstanding characteristics have motivated the scientific



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community to explore zirconia-based nanomaterials in a wide range of technological fields such as functional materials, *viz.*, catalysts,¹ sensors² and semiconductor devices,³ and structural materials, *viz.*, coating on cutting tools, ceramics,⁴ implants,⁵ *etc.* Apart from that, it can also be employed as a dielectric, electro-optic, and piezoelectric material due to its favourable optical and electrical properties.

Zirconia exists in three crystal phases, cubic, tetragonal, and monoclinic, in a typical air atmosphere and at distinct temperature ranges. The pure monoclinic phase is stable up to 1100 °C, the tetragonal phase is stable in the range of 1100–2370 °C and the cubic phase exists at temperatures over 2370 °C.⁶ Among the different phases, the cubic and tetragonal crystals are found to be unstable in bulk forms at ambient temperature. Therefore, considerable efforts have been dedicated to stabilizing the unstable cubic and tetragonal crystal phases by doping with several divalent or trivalent cationic stabilizers such as Ca²⁺, Mg²⁺, Sc³⁺, and Y³⁺ or by the reduction of particle/grain size to the nanometre scale.^{7–10} Moreover, the interconversion of phases such as cubic to tetragonal and subsequently tetragonal to monoclinic takes place at higher temperatures and also affects the volume of zirconia nanomaterials significantly which ultimately limits their applicability in different fields such as catalysis, coatings *etc.* Another limitation of zirconia nanoparticles has been their tendency to undergo agglomeration during their synthesis. Furthermore, limited methods have been reported for the synthesis of anisotropic zirconia nanoparticles with the desired size range. In this regard, it appears quite worthwhile to critically analyze the various methods, routes, and reaction conditions for the synthesis of such versatile nanomaterials. Therefore, the present review article contributes toward a comprehensive analysis of the state-of-art specific synthetic methods for the preparation of zirconia nanomaterials for the desired applications and provides a database for novel future research perspectives. There are two major sections in this review paper; the first section involves details about different synthesis methods of zirconia-based nanomaterials and a critical analysis of the effect of various reaction parameters and methods on size, shape, and morphology, and in the second part, a variety of applications of these materials have been included. Finally, perspectives and scope of further developments of such materials have been given.

1.1 Synthesis of zirconia nanomaterials

The research interest in synthesizing ZrO₂ nanoparticles by various methods has increased considerably in the last decade owing to their diverse applications. There have been interesting findings in terms of shape, size, morphology, yield, and purity when different methods were employed. Broadly, there are two approaches: the chemical method and greener synthesis which are discussed below:

1.1.1 Chemical methods. Zirconia nanoparticles are prepared by different chemical methods to obtain the desired shape and size, and crystallinity. Various chemical methods for the synthesis of zirconia-based nanomaterials include

microwave-assisted approach,¹¹ hydrothermal,¹² sol-gel,¹³ coprecipitation,¹⁴ and solution combustion method. A wide variety of zirconia precursors such as ZrOCl₂, ZrCl₄, and Zr(C₅H₇O₂)₄, and different reagents have been employed in these methods for the synthesis of nanomaterials of different shapes and sizes. Various reaction parameters such as zirconia salts, pH, reaction time, and temperature can also influence the morphology and size range of nanoparticles. The effect of different chemical synthesis methods and reaction conditions on the morphology and size range of ZrO₂ is given in Table 1. The comprehensive details of the chemical methods are given below.

(i) *Microwave-assisted approach.* Microwave-assisted approaches are quite popular as they are facile, advantageous, rapid, and energy efficient and hence, have been extensively employed in nanomaterial synthesis. In addition, microwave heating proves to be beneficial over conventional heating approaches in a way that microwaves could achieve elevated temperatures quickly in a pressure-controlled environment and exhibit a tendency to attain optimum crystallinity. The microwave technique depends on the effective heating of matter by microwave dielectric heating which is the capacity of a solvent and reagent to retain microwave energy and to convert it into heat. In this technique, heating is generally accomplished through dipolar polarization and ionic conduction.¹⁵ Different zirconia salts, *viz.* nitrate, sulphate, chloride, and chelating agents have been employed in this method. Singh and Nakate¹⁶ reported the synthesis of zirconia nanoparticles by using the microwave method at a low temperature (80 °C) and studied their photoluminescence (PL) properties. Powder X-ray diffraction analysis revealed the formation of combined phases of monoclinic and tetragonal ZrO₂. The microscopic analysis confirmed that the sizes of nanoparticles were less than 10 nm. The specific surface area of ZrO₂ NPs was found to be 65.85 m² g⁻¹. These particles exhibited excellent optical properties such as an optical band gap of 2.49 eV and an intense emission peak at 414 nm in the photoluminescence spectrum.

Recently, Yousaf *et al.*¹⁷ demonstrated the microwave synthesis of zirconia nanoparticles by using honey as a capping agent, for optical applications. The bandgap of the prepared tetragonal zirconia was found to be in the range of 4.7–4.82 eV. The structural characterization revealed the formation of single-phase tetragonal zirconium oxide with high purity under optimized conditions.

Similarly, Bukhari *et al.*¹⁸ prepared zirconia nanofibers by a microwave-assisted sol-gel approach. In this study, honey was used as a structure-directing and capping agent to enhance the stability, reduce the crystallite size of nanofibers, and prevent the hard agglomeration of nanoparticles. Consequently, soft or less agglomeration resulted in a stabilized tetragonal phase of zirconia by inhibiting its interphase transformation from tetragonal into monoclinic (Fig. 1).

The microwave wattage (100–900 W) also significantly affected the stabilization of zirconia nanoparticles. Powder X-ray diffraction study revealed the formation of highly pure tetragonal zirconia at low microwave power ~100 W with a smaller crystallite size of approximately 26 nm due to the



Table 1 Effect of different chemical synthesis methods and reaction conditions on the shape, size, and crystal structure of ZrO₂ NPs^a

| Chemical method | Materials | Solvent | pH | Reaction temperature/power and time | Thermal treatment | Crystal system, shape, size range, average size | Ref. |
|---------------------------|--|-------------------|--------|-------------------------------------|----------------------------------|---|----------|
| Microwave-assisted method | Zirconium oxychloride octahydrate, sodium hydroxide | — | — | 80 °C, 12 min | 400 °C, 2 h | Mixed phase (monoclinic and tetragonal), spherical (8–10) nm, 8.8 nm, — | 41 |
| | Zirconium oxychloride, anhydrous citric acid | — | 2 | —, 2 min | 450 °C, 4 h | Tetragonal, —, —, ~5–10 nm | 42 |
| | Zirconium oxychloride, sodium hydroxide | — | 12 | —, 6 min | 600 °C, 4 h | Monoclinic, —, 32–38 nm, — | 43 |
| | Zirconyl nitrate monohydrate, L-serine amino acid | Water | — | 800 W, 60 s | 400 °C, 600 °C, 800 °C, — | Cubic, —, —, 70, 76 and 77 nm | 44 |
| | Zirconium oxychloride octahydrate, Zirconium oxychloride octahydrate, honey | Water Water | — — | 100–1000 W, — ~100 W, — | —, — —, — | Tetragonal, —, —, — Tetragonal, —, —, ~26 nm | 19 18 |
| Hydrothermal method | Zirconyl oxynitrate | 1,4-Butanediol | — | 800 W, 10 min | —, — | Tetragonal, —, —, ~10 nm | 20 |
| | ZrO(NO ₃) ₂ ·xH ₂ O, sodium hydroxide | — | — | 150 °C, 4 h | 110 °C, 90 min | Monoclinic, —, -25 nm | 23 |
| | Zirconyl chloride octahydrate, glucose monohydrate | Water | — | 180 °C, 20 h | 550 °C, 4 h | —, sphere, -8 nm, — | 26 |
| | Commercial zirconia, sodium hydroxide | — | — | 150 °C, 85 h | —, — | Mixed phases (monoclinic, tetragonal and cubic), —, —, 15–30 nm | 12 |
| | Zirconium oxychloride, sodium hydroxide | Ethanol and water | 14 | —, — | 400 °C and 500 °C, 2 h and 0.5 h | Tetragonal, —, 2–11 and 3–12 nm, 4.5 and 6 nm | 21 |
| Sol-gel | Zirconyl nitrate, CTAB/SDS/triton, ammonium hydroxide | — | 12 | 90 °C, 12 h | 500 °C, 2 h | —, varied, varied, — | 27 |
| | Zirconium(IV) propoxide, acetic acid, nitric acid | Isopropanol | 2.5 | —, — | 700 °C, 1 h | Monoclinic, —, 25.39 nm, — | 32 |
| | Zirconium(IV) propoxide, nitric acid | — | — | —, 1 h | 500 °C, 2 h | Tetragonal, —, —, 30–60 nm | 33 |
| | Zirconium <i>n</i> -propoxide, glucose, fructose, aqueous ammonia | — | 9–10 | RT, 1 h | 700 °C, 3 h | Tetragonal and monoclinic, —, 10–30 nm, — | 13 |
| | Zr(SO ₄) ₂ ·H ₂ O, sodium lauryl sulfate | Benzyl alcohol | — | — | 600 °C, 5 h | Tetragonal and monoclinic, —, 9.7 nm, — | 45 |
| Co-precipitation | Zr(acac) ₄ , sucrose | Ethanol | ~12 | 110 °C, 10 min | 300–650 °C, — | Tetragonal and monoclinic, —, —, — | 35 |
| | Zirconium nitrate hexahydrate, potassium hydroxide | Water | >10 | RT, — | 900 °C, 4 h | —, spherical (193 nm), —, — | 38 |
| | Zirconyl chloride octahydrate, diammonium oxalate monohydrate, ethylene glycol | Water | — | Varied, varied | Varied, varied | Monoclinic, spherical, —, 12.85 nm | 39 |
| | Zirconium(IV) nitrate, tetraethyl <i>ortho</i> -silicate, ammonium hydroxide | Water and ethanol | 7 | 348 K, 10 min | 1173 K, 10 min | Tetragonal, —, 3.3 nm, — | 40 |
| | | | | | | | |

^a RT: room temperature.

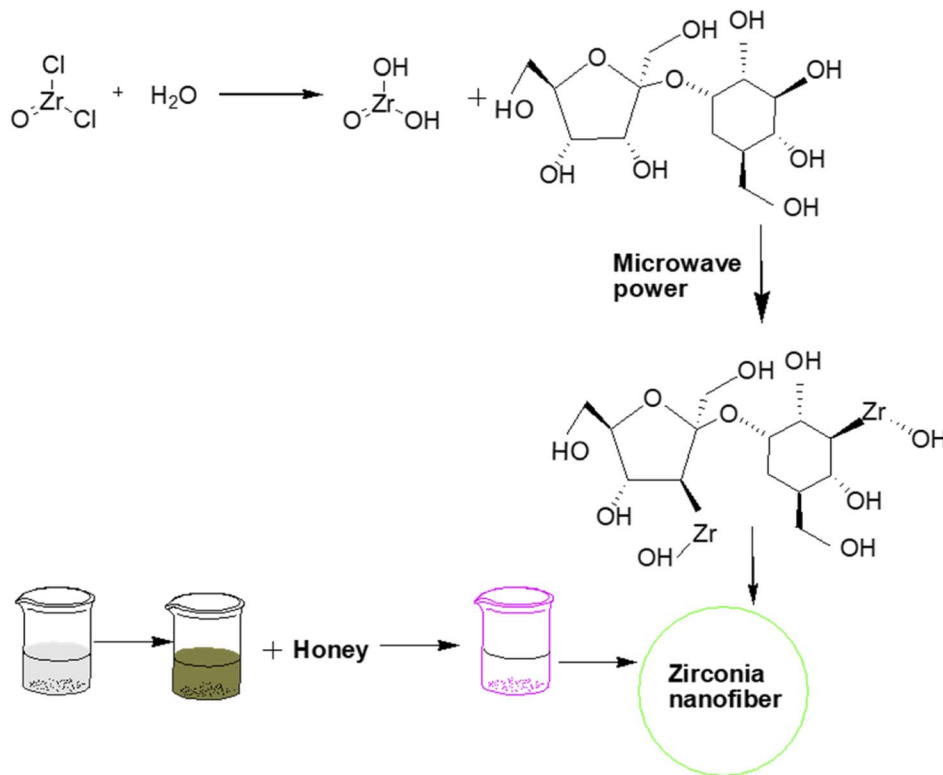


Fig. 1 Honey mediated microwave assisted sol-gel synthesis of stabilized zirconia nanofibers.¹⁸

presence of an optimum quantity of honey to coat the zirconia particles.

In another interesting report by Batool *et al.*,¹⁹ the preparation of zirconia nanoparticles and the dilemma of long-term stability were addressed. The obtained tetragonal zirconia nanoparticles were found to be stable even after a period of 6 to 12 months at room temperature.

Recently, Mishra *et al.*²⁰ demonstrated the synthesis of tetragonal zirconia nanoparticles having an average crystal size of 10 nm using a microwave-assisted solvothermal route. In this study, the bandgap of the prepared nanoparticles was found to be 3.67 eV.

(ii) *Hydrothermal and solvothermal approach.* The hydrothermal method is a promising approach for synthesizing nanomaterials and involves heating the mixture of precursors, appropriate agents (reducing, oxidizing, hydrolyzing), and suitable solvent in a closed vessel which facilitates the fabrication of nanocrystals with controlled surface alteration, morphology, size, and stability. Apart from that, such kinds of synthesis methods are also useful for doping heteroatoms in nanoparticles or fabrication of nanocomposites owing to the generation of high pressure in a closed vessel. These methods have been used for the preparation of zirconia nanoparticles as well as for the doping of zirconia with foreign species. For instance, Behbahani *et al.*¹² reported the hydrothermal synthesis of zirconia nanoparticles by using commercial monoclinic zirconia in an alkaline medium. Some of the parameters such as the concentration of the initial solution and

pH were varied accordingly to obtain novel nanoparticles with desired shape and size. All three phases of zirconia, *viz.*, cubic, tetragonal, and monoclinic phases, were observed and the corresponding sizes were found to be in the range of 15–30 nm.

Usually, attempting the synthesis of t-ZrO₂ nanoparticles by using an inorganic precursor without any surfactant results in the agglomeration of particles. Under such circumstances, the hydrothermal corrosion method, which involves a corroding medium such as H₂SO₄, HCl *etc.* to break down hard aggregates into dispersed fine nanoparticles during the hydrothermal treatment, could be employed. Zhou *et al.*²¹ demonstrated the hydrothermal-corrosion method to produce highly dispersed tetragonal zirconia nanoparticles (less than 10 nm) without using any surfactants. It was stated that the dispersity and crystallinity of the prepared tetragonal ZrO₂ powders can be modified by using reaction conditions such as calcination of the precipitates followed by hydrothermal HCl corrosion at different temperatures. In the follow-up study, Sagadevan *et al.*²² investigated the synthesis of spherical zirconia nanoparticles by the hydrothermal technique. The optical band gap of 5.02 eV and photoluminescence measurements confirmed the presence of oxygen vacancies accompanied by intrinsic defects in such NPs.

Ahmad *et al.*²³ successfully prepared ZrO₂ nanoparticles in an alkaline medium by employing the hydrothermal technique. Investigation of the structure and phase composition of the ZrO₂ nanoparticles by powder X-ray diffraction indicated the presence of pure monoclinic zirconia. The dielectric investigations suggested that ZrO₂ nanoparticles could be used for



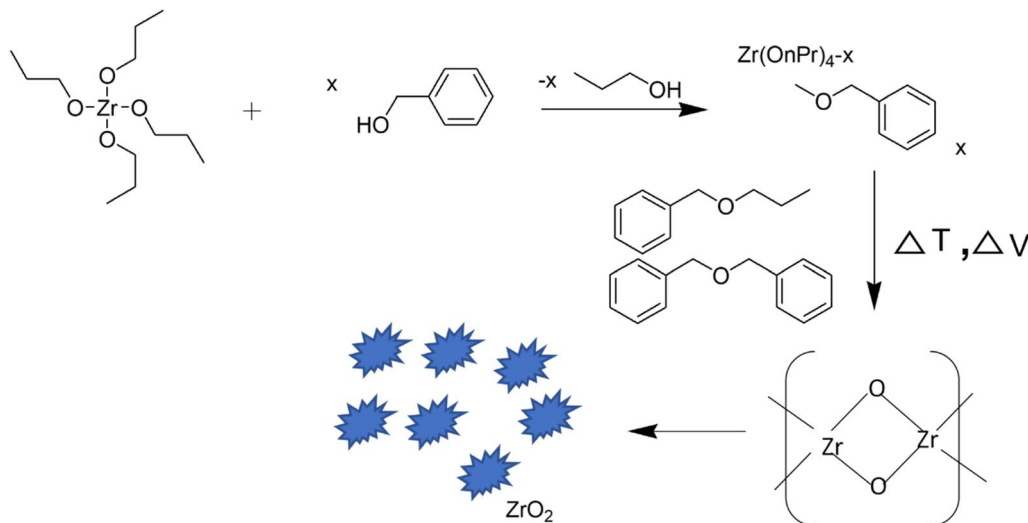


Fig. 3 Reaction mechanism of the nonaqueous synthesis of ZrO₂ nanoparticles.²⁸

method wherein, the sol can be obtained by dissolving a metal alkoxide or an organo-metallic precursor in water or alcohol which, under suitable reaction conditions (temperature, aging time, *etc.*), undergoes phase transformation into a gel-like structure. In the gel, owing to the network structure, polymeric particles remain separated, therefore the sizes of the particles can be controlled and agglomeration can be prevented. Finally, the gel upon suitable heat treatment could lead to the formation of nanoparticles. The principal advantage of the sol-gel method is the formation of uniform nanostructures at comparatively low temperatures as can be seen in Table 1. In this method, the reactivity of the precursors can be modified by alteration with chelation agents, consequently the gelation process could also be influenced, and ultimately the size and shape of the particles can also be modified. The different chelating reagents that have been employed for the preparation of nano zirconia are acetylacetonate, acetic acid, citric acid, and different sugars (*e.g.*, sucrose, maltose, glucose).

Madhusudhana *et al.*³² reported the synthesis of monoclinic zirconia nanoparticles by a facile sol-gel technique using acetic acid as a chelating agent. The resultant nanoparticles were found to be suitable for thermal applications due to their high resistance behavior to crack propagation. Lim *et al.*³³ developed a synthesis of tetragonal phase zirconia nanoparticles. Different concentrations of NH₃ solution and nitric acid were used with zirconium isopropoxide to analyze the impact on the crystallinity and particle size of ZrO₂. In contrast, smaller sizes of nanoparticles were obtained when nitric acid was used instead of ammonia in the reaction mixture.

Bahari *et al.*³⁴ investigated the preparation of ZrO₂ at a relatively lower temperature. It was observed that some amount of the monoclinic phase gets converted into metastable tetragonal ZrO₂ during heating. Heshmatpour *et al.*¹³ reported a simple sol-gel synthesis of zirconia nanoparticles of uniform size. Glucose and fructose were used as organic additives which could prevent the phase transition from the monoclinic to the tetragonal phase.

Likewise, the synthesis of zirconia thin films having a pure tetragonal phase was reported by Majedi *et al.*³⁵ via a sucrose-mediated sol-gel method. Initially, for preparing the sol, the zirconium(IV) acetylacetonate precursor was stirred in ethanol at 70 °C for 30 min followed by the addition of an ammoniacal solution as the hydrolyzing agent. Subsequently, sucrose was added to the solution as a gelation agent under stirring at 110 °C which results in the formation of the sol. The as-obtained sol was transformed into a xerogel by drying in an oven at 80 °C for 24 h. Furthermore, calcination at 350–500 °C results in t-zirconia, and calcination at 650 °C results in the formation of a mixture of the tetragonal and monoclinic phases. Such findings were expected as the post-calcination temperature could influence phase formation as well as phase transition.

Sponchia *et al.*³⁶ demonstrated a simple and efficient neutral surfactant-assisted sol-gel method for the synthesis of spherical mesoporous zirconia nanoparticles by using a zirconium propoxide precursor, surfactant, ethanol, and alkali halide, followed by solvothermal thermal treatment using ethanol and water. The obtained dried powder was subjected to heat treatment under vacuum conditions to extract the surfactant without any structural modification. The prepared material exhibited a high surface area in comparison to literature reports. The effect of the addition of different alkali halides on the shape and size of the nanoparticle was also investigated. The salt's solubility plays a crucial role in governing the particle size. Biological experiments demonstrated that mesoporous zirconia nanoparticles were biocompatible, degradable, and cell-permeable, therefore they could be employed for theranostic applications.

For instance, Davar *et al.*³⁷ synthesized pure monoclinic zirconia nanosheets by the polymeric sol-gel method. In this method, zirconium acetylacetonate was used as an organic metal precursor (Zr⁴⁺), and citric acid and ethylene glycol were used as the source of chelation and polymerization agents respectively. By varying the molar ratio of ethylene glycol to citric acid in the synthesis procedure, the phase as well as



morphology of the nanoparticles could be modified. With increase in mole ratio from 5 : 5 to 90 : 5, the phase changes from cubic to monoclinic, and morphology varies from semi-spherical to nanosheets.

(iv) *Co-precipitation technique.* Co-precipitation is an easy approach to prepare metal oxide nanoparticles using an aqueous solution of metal salts (chloride and nitrate metal salts) and bases such as ammonia or sodium hydroxide followed by heating or calcination at a suitable temperature. It involved nucleation followed by growth upon the addition of a precipitating agent; subsequently, heat treatment (calcination) led to the formation of metal oxide nanoparticles.

Ramachandran *et al.*³⁸ reported a co-precipitation method for the synthesis of zirconia nanoparticles using zirconium nitrate as the precursor and KOH as the precipitating agent. The ratio of zirconium nitrate and KOH (0.5, 1, and 1.5 M) was optimized to obtain crystalline zirconium oxide with spherical morphology. At the same line, Foo *et al.*³⁹ demonstrated the preparation of uniformly shaped zirconia nanoparticles having crystal sizes in the range of 6 to 35 nm. In this study, it was demonstrated that the calcination temperature in the reaction procedure could significantly impact the surface morphology, crystal size, and purity of the resulting ZrO₂ NPs. A higher calcination temperature increased the degree of agglomeration and size of crystals. Similarly, Huang *et al.*⁴⁰ also demonstrated that calcination temperature and time could drastically influence the crystallite size of tetragonal ZrO₂. In this study, when the freeze-dried precursor powder was calcined at 1173 K, crystallites of size 3.3 nm were obtained; however, comparatively bigger crystallites of 20 nm were observed upon calcination at 1473 K.

(v) *Solution combustion method.* The solution combustion method involves the propagation of self-sustained and coupled complex exothermic reactions in an aqueous solution which could result in the formation of a ZrO₂ powder along with emission of various gases such as N₂, CO, CO₂, *etc.* It starts with the burning/oxidation of a fuel which is an organic compound employed as a reactant along with a metal precursor. This method has been used in the synthesis of a variety of nanomaterials, including zirconia nanoparticles. Manjunatha *et al.*⁴⁴ synthesized the pure cubic phase of ZrO₂ mesoporous nanoparticles *via* a microwave-assisted combustion method. In this procedure, zirconyl nitrate monohydrate (ZrO(NO₃)₂·H₂O) was used as the metal precursor and L-serine amino acid was employed as the fuel. Both components were dissolved in a minimum quantity of de-ionized water. Subsequently, the excess water was allowed to evaporate to obtain a viscous gel which upon microwave irradiation at 800 W for 60 seconds led to vigorous boiling followed by dehydration. Subsequently, the material gets burnt leading to the liberation of gases such as CO₂, N₂, and water vapour, and the formation of a white ZrO₂ powder having a cubic phase. This study indicated that both the crystallinity and band gap (~5.21 eV) of the prepared nanoparticles increases with increase in calcination temperature. An increase in band gap was observed due to a decline in the number of intermediate energy levels owing to an increase in orderliness with a smaller number of defects. The

photoluminescence spectrum displayed an intense violet emission band peak centred at 387 nm, owing to the existence of oxygen vacancies in the zirconia matrix. The surface area analysis revealed that such nanomaterials were mesoporous in nature and the pore size also tends to decrease with an increase in the calcination temperature.

Dwivedi *et al.*⁴² prepared ZrO₂ NPs by a microwave-driven citrate sol-gel technique. In this study, the prepared zirconia nanoparticles exhibited a smaller size due to gelation and rapid combustion.

Prakash Babu *et al.*⁴⁶ also introduced a simple low-temperature solution combustion method without any stabilizing agent and calcination. In this method, zirconyl nitrate as the metal precursor and oxalyl di-hydrazide as the fuel were dissolved in water. The aqueous solution was taken in a dish and subjected to heat treatment at 400 °C in a preheated muffle furnace. Zirconia nano-powder having a cubic phase and crystallite size in the range from 6 to 12 nm was obtained.

Similarly, Tang *et al.*⁴⁷ reported the synthesis of hollow mesoporous zirconia nanocapsules with a porous shell structure. In this study, first, monodispersed silica nanoparticles were prepared *via* Stober's method using concentrated ammonia, water, ethanol, and TEOS followed by the formation of the SiO₂@ZrO₂ composite using Zr(OBu)₄. Finally, the composite was treated with NaOH (5 M) to remove silica which resulted in the formation of hexagonal mesoporous ZrO₂.

Though different methods for the formation of specific phases have been reported, microwave-assisted sol-gel and solution combustion methods can be identified as promising choices for the formation of specific phases. Particularly, with the solution combustion method, formation of the cubic phase has been observed predominantly. This could be due to the attainment of high temperature that favors the formation of the cubic phase which is also a high-temperature stable form of zirconia. Interestingly, the use of microwave treatment in the case of the sol-gel has also resulted in the formation of selective phases, particularly the tetragonal phase. Apart from that, employing a corrosive solvent (HCl, H₂SO₄) during the hydrothermal treatment of zirconia nanoparticles could be a useful approach for obtaining monodisperse nanoparticles.

In most of the published literature studies on zirconia nanomaterials, the formation of spherical particles has been reported. However, in some studies, the formation of unique shapes has also been reported using unusual reaction conditions and precursors. For instance, the formation of nanosheets has been reported *via* the sol-gel method using zirconium acetylacetonate as the metal precursor, and the formation of cubic and star-like shapes has been reported *via* the hydrothermal method at a comparatively higher temperature of 240–250 °C.^{29,30,48} Interestingly, comparatively smaller particle sizes have been observed along with the pure tetragonal phase in all types of chemical synthesis methods.

1.1.2 Greener methods. Nowadays greener pathways for the synthesis of nanoparticles are being explored as they could enhance their biocompatibility, reduce the cost of the designed nanomaterial, and take care of environmental concerns.⁴⁹ In this context, the incredible commitment of Paul and Richard



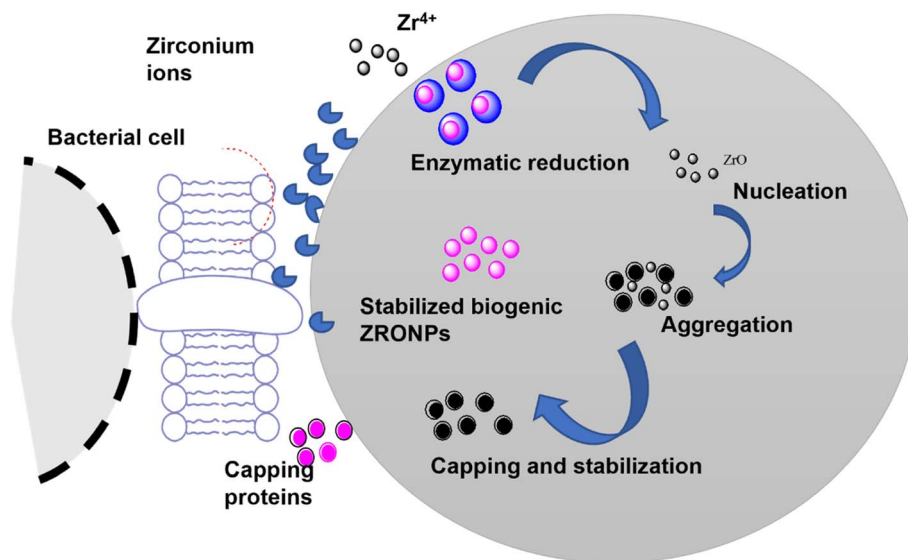


Fig. 4 Schematic representation of proposed extracellular biosynthesis mechanisms of ZrO_2 NPs. Nitrate reductase enzyme is responsible for the reduction of metal ions to their respective metal atoms that lead to nucleation and formation of NPs.⁵³

toward making nanotechnology and green science doable is appreciated. Synthesis of metal nanoparticles using renewable and natural sources of reducing, stabilizing, and capping agents (microbes and parts of plants) is considered a greener approach. Microbial synthesis of nanoparticles includes microorganisms, such as bacteria, fungus, *etc.*, which could play a significant role in the reduction and stabilization of nanoparticles.^{50,51} Nanoparticles can be synthesized *via* both intracellular and extracellular pathways. The intracellular pathway involves the transportation of metal ions into the cell by interaction with the negative receptors of the cell wall followed by the reduction of metallic species by the enzymes of the cell, whereas the extracellular synthesis of the nanoparticle is mediated by nitrate reductase which reduces the zirconium ions. Subsequently, the formation of nanoparticles takes place *via* nucleation, aggregation, and growth (Fig. 4).^{52,53} Furthermore, stabilization of nanoparticles is known to occur due to capping agents (protein species) released by the cell wall. Uddin *et al.*⁵⁴ reported extracellular biosynthesis of ZrO_2 nanoparticles by using potassium hexafluorozirconate (K_2ZrF_6) as a precursor and the fungus *Humicola* sp. The characterization results revealed the presence of protein on the surface of quasi-spherical NPs. In the follow-up study, Bansal *et al.*⁵⁵ demonstrated the synthesis of zirconia nanoparticles *via* extra-cellular hydrolysis

of the metal halide precursor (ZrF_6^{2-}) by the cationic protein species (molecular weight 24 and 28 kDa) of the fungus *Fusarium oxysporum*. It involves the formation of a complex at room temperature under acidic conditions. Shanmugasundaram *et al.*⁵⁶ demonstrated a one-pot biosynthesis of non-toxic ZrO_2 nanoparticles (of size 44 ± 7 nm) using *Acinetobacter* sp. KCSI1 at room temperature. The zeta potential of the as-synthesized zirconia nanoparticles was found to be 36.5 ± 5.46 mV. Similarly, Ahmed *et al.*⁵³ demonstrated the synthesis of spherical-shaped zirconia nanoparticles (33–75 nm) by using *Enterobacter* sp. The key reaction parameters and important findings from such reports have been summarized in Table 2.

Nowadays, extracts of different parts of plants such as roots, leaves, flowers, and fruits are also being used for the preparation of nanomaterials. Various phytochemicals, *viz.* flavones, terpenoids, sugars, ketones, aldehydes, carboxylic acids, and amides, are found to be responsible for the reduction of metal precursors. These species act as both stabilizing and reducing agents. In general, the plant-mediated synthesis of zirconia NPs is a simple and facile process, in which a zirconia salt, such as zirconium oxychloride or zirconia nitrate is mixed and stirred with an appropriate plant extract at a suitable temperature followed by heat treatment. However, there are limited reports on the mechanism of formation and stabilization of zirconia

Table 2 Effect of different microorganisms and pH on the shape, size, and crystal structure of ZrO_2 NPs

| Microorganism | Zirconia source | pH | Size range, average size (nm) | Crystal system, morphology | Ref. |
|---|-----------------------------------|-----|-------------------------------|------------------------------|------|
| <i>Humicola</i> sp. (fungus) | Potassium hexafluoro zirconate | 9 | 13 nm | –, quasi-spherical | 54 |
| <i>Fusarium oxysporum</i> (fungus) | Potassium hexafluoro zirconate | 3.6 | 7.3 ± 2.0 nm | Monoclinic, – | 55 |
| <i>Acinetobacter</i> sp. KCSI1 (bacteria) | Zirconium oxychloride octahydrate | 2 | 44 ± 7 nm | – | 56 |
| <i>Pseudomonas aeruginosa</i> (bacteria) | Zirconium oxychloride octahydrate | – | 6.41 nm | Monoclinic and tetragonal, – | 66 |
| <i>Enterobacter</i> sp. (bacteria) | Zirconium oxychloride octahydrate | – | 33–75 nm | –, spherical | 53 |



NPs from such biomolecules. Different reaction parameters, and resultantly different findings in terms of crystal structure and size distribution from recent reports on plant-mediated synthesis have been summarized in Table 3.

Silva *et al.*⁵⁷ demonstrated a green approach for the preparation of zirconia nanoparticles, exhibiting monoclinic and tetragonal phases by using *Euclea natalensis* plant extract. In this study, factorial design suggested that a higher concentration of precursor, lower amount of extract, and higher calcination temperature supported the synthesis of the tetragonal phase. Moreover, a possible mechanism (Fig. 5) for the preparation of ZrO₂ NPs using phytochemicals of *Euclea natalensis* such as pentacyclic terpenoids (lupeol, betulin, and β -sitosterol) and naphthoquinones (diospirine, shinanolone, and 7-methyl juglone) is also given in this report. However, there is ambiguity in the mechanism of formation of ZrO₂ NPs. It involves three steps: the reduction of zirconyl ions, the grouping of nanoparticles, and the growth of nanoparticles. But, only the reduction of zirconium ions could not lead to the formation of ZrO₂ as oxidation is also a necessary step for its formation. In another report,⁵⁸ bioreduction of zirconium ions, followed by complexation with water molecules, chelation with phytochemicals and finally oxidation were the major suggested steps in the plant-mediated synthesis of zirconia. As per our recent investigation on the plant-mediated synthesis of ZrO₂, which is in the process of publication, a change in the oxidation state of zirconium from +4 to +2 followed by complexation with phyto-components was observed. The complex upon heat treatment regains its oxidation state of +4 in zirconia as that of the precursor (ZrOCl₂). These findings were supported by DFT calculations.

Another study on the green fabrication of ZrO₂NPs was reported by Jalill *et al.*⁵⁹ and the prepared nanoparticles exhibited excellent antifungal and antibacterial activity. Shanthi *et al.*⁶⁰ synthesized ZrO₂ nanoparticles by employing aqueous leaf extracts of *Acalypha indica* as a reducing agent. Encouraged by previous studies, Padma *et al.*⁶¹ fabricated zirconium dioxide nanoparticles using *Azadirachta indica* leaves. Similarly, Sathish *et al.*⁶² synthesized zirconia nanoparticles using *Curcuma longa* tuber. In this study, a higher pH was found to be favorable for the fabrication of zirconia NPs. Similarly, Gowri *et al.*⁶³ investigated the greener synthesis of spherical zirconia nanoparticles using *Aloe vera* plant extract. The nanoparticles were found to have sizes in the range of 50–100 nm. Majedi *et al.*⁶⁴ reported the synthesis of cubic phase zirconia nanoparticles using zirconium acetate as the precursor and lemon juice as the reducing agent. The impacts of sucrose addition on enhancing the particle size and the yield of the product were examined. A mixture of 20 mL lemon juice and sucrose was found to show a positive effect on uniform morphology having an average particle size of 21 nm. The as-prepared nanoparticles were employed in electrolyte materials for intermediate or low-temperature solid oxide fuel cells. Saraswathi *et al.*⁶⁵ demonstrated the fabrication of tetragonal zirconia nanoparticles (2 nm) using *Lagerstroemia speciosa* leaves. The authors investigated the photocatalytic activity of the nanoparticles by

Table 3 Effect of different reaction conditions and plant extract on the shape, size, and crystal structure of ZrO₂ NPs

| Plant | Zirconia source | Conc. of plant extract, precursor, ratio | Reaction temperature, time | Thermal treatment (drying and calcination) | Size range, average size (nm) | Crystal system, shape | Ref. |
|--------------------------------------|-----------------------------------|--|----------------------------|--|-------------------------------|--------------------------------------|------|
| <i>Euclea natalensis</i> roots | Zirconyl chloride octahydrate | 50 g L ⁻¹ , 0.03 mol L ⁻¹ , 1 : 1 | –, 3 h | 105 °C for 24 h, 550 °C for 3 h | 5.25–41.71 nm, – | Monoclinic and tetragonal, spherical | 57 |
| <i>Acalypha indica</i> leaves | Zirconyl chloride octahydrate | 100 g L ⁻¹ , 0.1 mol L ⁻¹ , 1 : 5 | 80 °C, 2 h | 200 °C for 2 h, – | 58–72 nm, – | Monoclinic, – | 60 |
| <i>Azadirachta indica</i> leaves | Zirconyl chloride octahydrate | 100 g L ⁻¹ , 50 mol L ⁻¹ , – | 50–60 °C, – | 120 °C, 600 °C for 5–6 h | –, – | –, – | 61 |
| <i>Aloe vera</i> leaves | Zirconium oxychloride octahydrate | 500 g L ⁻¹ , 0.4 mol L ⁻¹ , – | Room temperature, 4 h | –, 500 °C | 27.4 nm, 50 nm | Tetragonal, spherical | 67 |
| <i>Eucalyptus globulus</i> leaves | Zirconium oxychloride | 100 g L ⁻¹ , 0.0098 mol L ⁻¹ , 1 : 1 | –, 2–3 h | 80 °C, 600 °C for 3 h | 9–11 nm, – | Monoclinic and cubic, spherical | 68 |
| <i>Lagerstroemia speciosa</i> leaves | Zirconium nitrate | 150 g L ⁻¹ , 0.2 N, 1 : 1 | 90 °C, 3 h | –, – | –, 56.8 nm | Tetragonal, – | 65 |
| <i>Curcuma longa</i> tuber | Potassium hexafluoro zirconate | 25 g L ⁻¹ , 0.00098 mol L ⁻¹ , – | 25 °C, 72 h | –, – | –, 41–45 nm | Monoclinic, – | 62 |
| <i>Citrus aurantiifolia</i> fruit | Zirconium(IV) acetate | 10–30 mL, 0.038 mol L ⁻¹ , – | 90 °C, 3 h | –, 750 °C for 2 h | 21 nm, 20 nm | Cubic, – | 69 |





Table 4 Zirconia-based heterogeneous catalysts for transesterification and esterification

| Reactants | Catalyst | Methanol to oil molar ratio | Reaction temperature (°C) and time | Catalyst loading | Type of reaction | Conversion/yield | Ref. |
|---|---|--------------------------------|--------------------------------------|------------------|-------------------------------------|----------------------------|------|
| Waste cooking oil and methanol | Sr/ZrO ₂ | 29 : 1 | 115.5 °C, 169 min | 2.7% | Transesterification, esterification | 79.7% | 70 |
| Soybean oil and methanol | C ₄ H ₄ O ₆ HK/ZrO ₂ | 16 : 1 | 60 °C, 2.0 h | 6% | Transesterification | 98.03% | 71 |
| Oleic acid and methanol | Chlorosulfonic acid modified zirconia | 8 : 1 | 100 °C, 12 h | 3% | Esterification | ~100% | 73 |
| Refined soybean oil and methanol | CdO/ZrO ₂ | 40 : 1 | 135 °C, 3 h | 7% | Transesterification | 97% | 75 |
| Silybum marianum oil and methanol | KOH/ZrO ₂ | 15 : 1 | 60 °C, 2 h | 6% | Transesterification | 90.8% | 74 |
| Tannery waste sheep fat and methanol | Ferric-manganese-doped sulfated zirconia | 15 : 1 | 65 °C, 300 min | 8% | Transesterification | 98.7% | 76 |
| Oleic acid and jatropha oil | SO ₄ ²⁻ /Zr-KIT-6 | 20 : 1 | 120 °C, 6 h | 4% | Esterification, transesterification | 95%, 85% | 78 |
| Soybean oil and methanol | Mesoporous sulfated zirconia | 12 : 1 | 120 °C, 4 h | 4% | Transesterification | 94.9% | 79 |
| Soybean oil and methanol | Sulfated zirconia | 20 : 1 | 150 °C, 6 h | 3% | Transesterification | 84.6% | 80 |
| Sunflower oil | Sulphated ZrO ₂ -MCM-41 | 9 : 1 | 60 °C, 30 min | 5% | Esterification | 96.9% | 81 |
| Soybean oil and methanol | ZrO ₂ -supported bamboo leaf ash (ZrO ₂ /BLA) | 15 : 1 | -, 30 min | 12% | Transesterification | 92.75% | 82 |
| Oleic acid and methanol | SO ₃ H@ZrP | 9 : 1 | 65 °C, - | 5.0% | Esterification | 89% | 95 |
| <i>Nannochloropsis</i> sp. (lipid) and methanol | Bi ₂ O ₃ /ZrO ₂ (CTAB) | 90 : 1 | 80 °C, 6 h | 20% | Transesterification, esterification | 73.21% | 83 |
| Soybean oil/tributyrin, and methanol | UiO-66 | 44 : 1.1/52 : 1.3 respectively | 140 °C, 5 h/120 °C, 5 h respectively | 0.1 g | Transesterification | 98.5%, 99.3% | 90 |
| Oleic acid and methanol | UiO-66(Zr)-green, UiO-66(Zr)-NO ₂ -green, and UiO-66(Zr)-NH ₂ | 39 : 1 | 60 °C, 4 h | 6% | Esterification | 86%, 90%, 97% respectively | 91 |
| Soybean oil and methanol | AlLs/POM/UiO-66-2COOH | 35 : 1 | 110 °C, 6 h | 10% | Esterification, transesterification | 95.8% | 92 |
| Tributyrin and methanol | Zr-MOFs-S | 52 : 1.3 | 140 °C, 6 h | 0.1 g | Transesterification | — | 93 |
| Oleic acid with methanol | ZrSiW/UiO-66 | 20 : 1 | 150 °C, 4 h | 0.24 g | Esterification | 98.0% | 94 |

heterogeneous catalyst in contrast to sulfated zirconia and it could also be employed for a pre-treatment step of esterification of oleic acid; consequently, saponification and catalyst inactivation were avoided. Similarly, Takase *et al.*⁷⁴ investigated the transesterification of a non-edible oil (*Silybum marianum*) using KOH-impregnated zirconia NPs as a catalyst. Patil *et al.*⁷⁵ prepared zirconia-supported cadmium oxide *via* a co-precipitation approach and used it for the transesterification of soybean oil with methanol. The activation energy and yield of the reaction were found to be $41.18 \text{ kJ mol}^{-1}$ and 97% respectively under optimum conditions. Booramurthy *et al.*⁷⁶ investigated a sulfated zirconia nanocatalyst for transesterification using tannery waste sheep fat. Lin *et al.*⁷⁷ demonstrated the aqueous phase synthesis of ZrO_2 nanoparticles (5–10 nm) having Lewis acidic sites and employed it for the preparation of biofuel from jatropha oil. In this study, it was demonstrated that ZrO_2 nanoparticles have a higher capacity for phospholipid adsorption in contrast to commercial ZrO_2 powder owing to their higher surface area. Gopinath *et al.*⁷⁸ demonstrated the preparation of mesoporous sulfated Zr-KIT-6 *via* a hydrothermal route and employed it as a catalyst for the production of biodiesel. In this study, biodiesel was prepared by the esterification of oleic acid oil and transesterification of jatropha oil. This catalyst was significantly acidic due to the presence of Zr and sulphation, which resulted in an increase in activity for both transesterification as well as esterification. The prepared catalyst showed good stability and reusability and could be recycled up to three times. Similarly, Luo *et al.*⁷⁹ investigated the suitability of mesoporous sulphated zirconia NPs as an acidic catalyst to prepare biodiesel. In this study, the sulfated zirconia nanomaterial was prepared by a one-step liquid-crystal template method, where sodium dodecyl sulfate acted as a template and

sulfonating agent which improved the catalytic performance for transesterification. It was observed that higher calcination temperature resulted in the loss of catalytic activity such as loss of sulphate species, significant loss of acid sites, and the increased transformation of tetragonal zirconia into monoclinic zirconia, and reduced the surface area. Shi *et al.*⁸⁰ prepared the sulfated t-zirconia catalyst having super acidic sites by directly impregnating ammonium persulphate followed by calcination at $500 \text{ }^\circ\text{C}$. It was employed as a catalyst for the preparation of biodiesel and demonstrated excellent yield (84.6%). Dehghani *et al.*⁸¹ prepared a nanocatalyst using nano-sulfated zirconia supported over MCM-41 *via* an ultrasound-assisted impregnation/hydrothermal hybrid approach. The prepared catalyst showcased excellent activity to produce biodiesel (96.9%). Fatimah *et al.*⁸² reported a heterogeneous ZrO_2 catalyst using bamboo leaf ash as a support for transesterification. Rahman *et al.*⁸³ investigated the zirconia based bifunctional heterogeneous nano catalysts $\text{ZrO}_2(\text{P123})$ and $\text{ZrO}_2(\text{CTAB})$ for the conversion of microalgae to biodiesel. Surfactant-assisted sol-gel method was followed to obtain high surface area catalyst. The introduction of Bi_2O_3 on zirconia enhanced the density of total acidic and basic sites of the catalyst. A possible mechanism for the simultaneous reactions using the bifunctional acid-base zirconia catalyst for biodiesel production is depicted in Fig. 6.

Metal organic frameworks (MOFs) constitute a new emerging class of porous materials. However, these materials are still in early stages of development and gradually progressing towards diverse applications such as catalysis, drug delivery, fluorescence sensing, *etc.* Zr based metal organic frameworks (Zr-MOFs) are most likely considered as encouraging MOF materials owing to their excellent physico-chemical properties and outstanding chemical stability. In 2008, an extraordinary

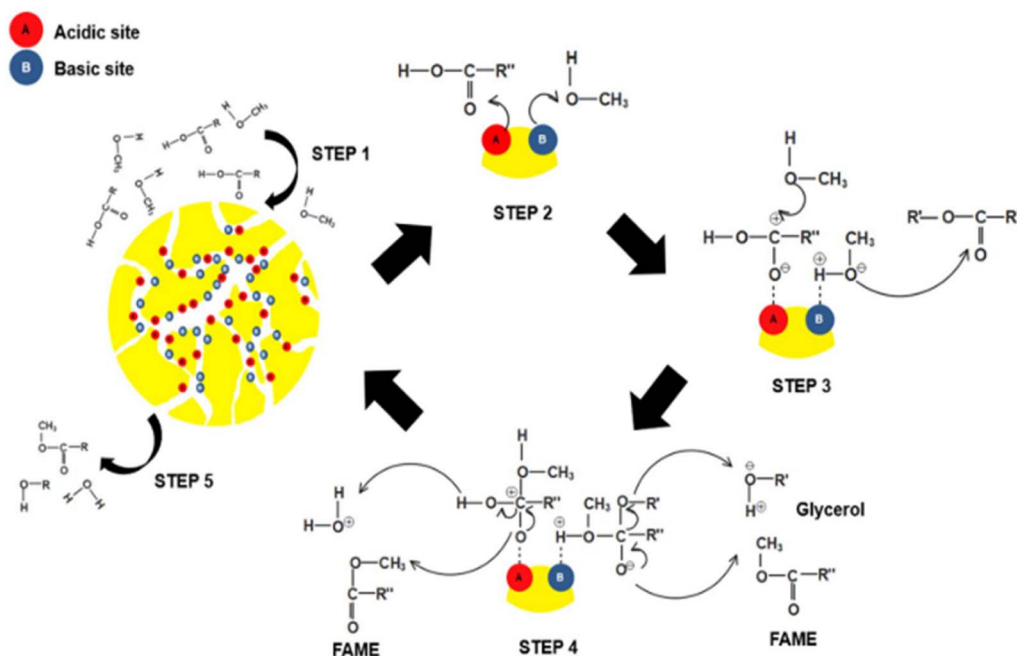


Fig. 6 The reaction mechanism of biodiesel production using a bifunctional acid–base catalyst.⁸³



discovery was made by Cavka *et al.*,⁸⁴ wherein 12-coordinated $Zr_6(\mu_3-O)_4(\mu_3-OH)_4(CO_2)_{12}$ clusters were synthesized, commonly referred to as UiO-66. It was found that the chemical and hydrothermal stability of UiO-66 was much higher than that of the other reported MOFs.^{85–88} Since then, extensive efforts have been dedicated for the synthesis of Zr based MOFs, and the investigation of their properties, structures, functions and applications has been carried out intensively.

Schaate *et al.*⁸⁹ described the very first example of employing a modulated fabrication approach to prepare Zr based MOFs and studied the effects of different modulators such as acetic acid, water and benzoic acid on the formation of Zr-MOFs. Interestingly, it was found out that among all modulators, only benzoic acid affected the synthesis process; as its concentration was increased, bigger and quite distinct crystals were expected to form from the intermediates of UiO-66. Hence, the first large single crystals of Zr-MOF were obtained by this strategy, as confirmed by the single-crystal XRD technique. An interesting study by Zhou *et al.*⁹⁰ revealed the catalytic activity of UiO-66 prepared by using terephthalic acid and $ZrCl_4$. The catalyst demonstrated excellent recyclability for the preparation of biodiesel using tributyrin and soybean oil with methanol. The performance of the catalyst was significantly dependent upon linker defects under different reaction parameters. Similarly, Abou-Elyazed *et al.*⁹¹ reported the use of UiO-66(Zr)-green, UiO-66(Zr)- NO_2 -green, and UiO-66(Zr)- NH_2 green for the preparation of biodiesel. Among all these MOFs, UiO-66(Zr)- NH_2 is found to be an efficient catalyst with an activation energy of $15.13 \text{ kJ mol}^{-1}$. Xie *et al.*⁹² demonstrated a one-pot preparation of biodiesel from low-cost acidic vegetable oils by using the AILS/POM/UiO-66-2COOH catalyst with 95.8% conversion of oil under optimized conditions. The obtained catalyst exhibited various merits such as good activity, higher surface area, reusability up to 5 cycles, and tolerance to high amounts of free fatty

acid (9 wt%) and water (3 wt%) in the feedstock. Lu *et al.*⁹³ synthesized sulphated Zr-based MOFs which had flower-like mesoporous nanosheets of zirconium and sulphate. In this study, the catalyst exhibited excellent activity, larger surface area ($186.1 \text{ m}^2 \text{ g}^{-1}$), and recyclability for the transesterification reaction. Recently Zhang *et al.*⁹⁴ prepared a ZrSiW/UiO-66 metal-organic framework by a hydrothermal method and used it for the biodiesel production. The prepared catalyst exhibited good activity, high surface area, and showcased heterogeneous nature as per the mechanism depicted in Fig. 7.

It is manifested from the literature cited in Table 4 that zirconia-based nanomaterials with enhanced acidity (sulphated zirconia) have been investigated extensively as a catalyst for simultaneous transesterification and esterification. However, a comparatively higher temperature and molar ratio of methanol to oil is required for the high yield of the reaction in contrast to the catalyst with enhanced basic sites (zirconia modified with KOH). Therefore, exploring Zr-based MOF and composites with basic sites could result in an efficient catalyst for transesterification.

1.2.1.1 Photocatalysis. ZrO_2 , which is a semiconductor having a broad direct bandgap and high negative value of the conduction band, has also been used as a photocatalyst due to its suitable thermal stability, and optical and electrical properties. It has been used as a photocatalyst for the degradation of organic water pollutants such as antibiotics and dyes. In Table 5, key parameters for the photocatalytic reactions using ZrO_2 -based photocatalysts have been summarized. As can be seen in Table 5, there is variation in the photocatalytic activity of zirconia even for the same dye owing to different particle sizes and consequent different band gaps. Therefore, lights from both UV and visible region have been employed with ZrO_2 for the photocatalytic reactions. For instance, Ajabshir *et al.*⁹⁶ demonstrated the photocatalytic oxidation of the Eriochrome

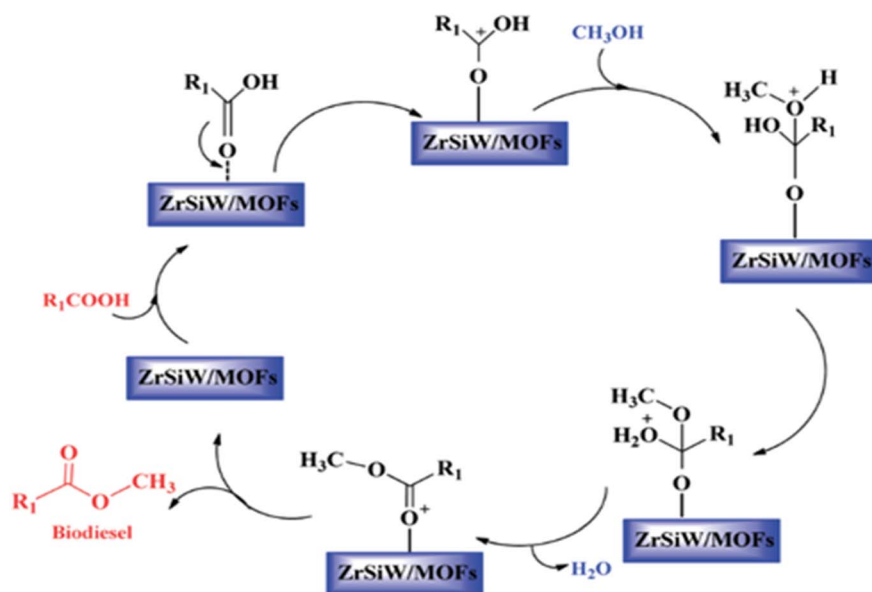


Fig. 7 Possible mechanism of esterification.⁹⁴



Black T dye using sphere-like ZrO₂ NPs under UV light. On the other hand, star-like ZrO₂ NPs exhibited selective photocatalytic activity for the degradation of rhodamine B, methyl orange, and Congo red under visible light.³⁰ The mesoporous form of zirconia also demonstrated photocatalytic activity under the illumination of visible light for the degradation of tetracycline.⁹⁷ The $\cdot\text{O}_2^-$ and h^+ radicals were found to be accountable for the degradation of tetracycline. In another interesting report, Basahel *et al.*⁹⁸ proposed a mechanism (Fig. 8) for the degradation of methyl orange using zirconia nanostructures in varied crystalline forms, *viz.*, monoclinic, tetragonal, and cubic. The monoclinic zirconia showed relatively higher photocatalytic activity than the other phases (tetragonal and cubic) due to the high crystallinity, large pore size distribution, oxygen-deficient zirconium oxide phase, and high density of surface hydroxyl groups. It involved the adsorption of methyl orange onto the surface of the nanoparticles. The exposure to UV light resulted in the formation of electron-hole pairs which react with the dissolved oxygen molecules and surface hydroxyl groups to form superoxide anion radicals and reactive hydroxyl radicals respectively.

As can be seen in Table 5, different sizes of ZrO₂ NPs and their nanocomposites have been used for the degradation of different dyes. Therefore, a clear conclusion about their photocatalytic activity can't be drawn directly. However, a generalization on the band-gap which is responsible for photocatalysis can be made. The lowest band gap of ZrO₂ was observed in its composites with reduced graphene oxide followed by that for the pure ZrO₂ NPs having a spherical shape and an average size of 17 nm. Thus, composites of small size ZrO₂ could further result in promising photocatalytic activities.

1.2.2 Anti-bacterial and anti-cancer activity. The use of ZrO₂-based nanomaterials in some biomedical solutions such as dentistry and drug delivery has been identified as an innovative approach owing to their low cost and less toxic nature. Therefore various biological properties, *viz.* antibacterial, antifungal, anticancer, and antioxidant, of zirconia-based structures have been investigated.^{47,65,67,68} The important properties and key advancements are summarized below.

Gowri *et al.*⁶⁷ synthesized zirconia nanoparticles of 50–100 nm using leaves of aloe vera and investigated their antimicrobial and antifungal properties. Abdul *et al.*⁵⁹ demonstrated excellent antifungal and antibacterial properties of ZrO₂ NPs against *F. moniliforme*, *F. graminearum*, *E. coli*, and *S. aureus*. Similarly, Imran *et al.*¹⁰⁴ prepared Fe₃O₄-doped ZrO₂ nanoparticles of average size ~23 nm by the microwave method and investigated their antimicrobial activity. These nanoparticles (NPs) were used as dental fillers. Kumaresan *et al.*¹⁰⁵ demonstrated the antibacterial activity of zirconia NPs against *Bacillus subtilis*, *Escherichia coli*, and *Salmonella typhi*. The authors used marine brown alga, *Sargassum wightii* for the synthesis of zirconia nanoparticles having the size of ~4.8 nm. Masim *et al.*¹⁰⁶ prepared a polyaniline zirconia (PANI-ZrO₂) composite and employed it for the treatment of wastewater. The obtained composite showed antibacterial and anti-corrosion activity and acted as a phosphate adsorbent material.

Table 5 Photocatalytic activities of greener ZrO₂-based nanomaterials

| Nanomaterial | Band gap (eV) | Average size (nm), shape, size range (nm) | Light source | Dye | Time of irradiation (min) | Degradation (%) | Ref. |
|---|----------------------------|--|--------------|----------------------------------|---------------------------|-----------------|------|
| ZrO ₂ | 3.78 | 17, spherical, – | Sunlight | RY 160 | 120 | 94 | 99 |
| ZrO ₂ | 4.9 | 15, spherical, 10–18 | UV light | Methylene blue and methyl orange | 240 | 91, 69 | 100 |
| ZrO ₂ | – | 56.8, oval, – | Sunlight | Methyl orange | 290 | 94.58 | 65 |
| ZnO-ZrO ₂ | 3.96, 3.99, 3.97, and 4.01 | 79.56, 98.78, 54.86, and 67.43, –, – | Solar light | Rhodamine 6G | 330 | 98.94 | 101 |
| Mg-doped ZrO ₂ | Varied | –, aggregated spherical, tetragonal and hexagonal, – | UV light | Rhodamine B | 60 | 93 | 8 |
| Reduced graphene oxide-ZrO ₂ | 3.25 | –, –, – | Sunlight | Reactive Blue 4 | 120 | 93 | 102 |
| V ₂ O ₅ /ZrO ₂ | 3.93 | 34–50 | Solar light | Methyl orange and picloram | 75 | 76.94% and 86% | 103 |



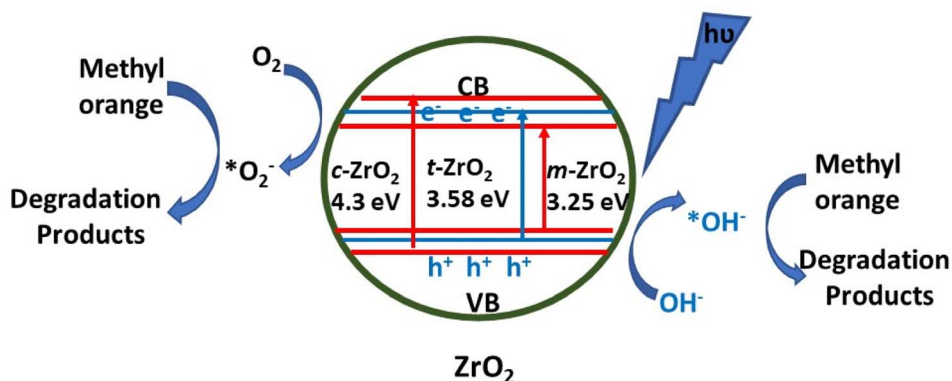


Fig. 8 Schematic representation of the processes leading to photocatalytic degradation.⁹⁸

Fathima *et al.*¹⁰⁷ reported the fabrication of zirconia nanoparticles having sizes in the range of 15–21 nm and investigated their antimicrobial activity. In this study, they identified antimicrobial activity against Gram-positive (*Bacillus subtilis* and *Staphylococcus aureus*) and Gram-negative bacteria (*Escherichia coli* and *Pseudomonas aeruginosa*), and found superior inhibitory activity against *Pseudomonas aeruginosa* at the concentration of 100 $\mu\text{g mL}^{-1}$. The antimicrobial activity of ZrO_2 NPs for different bacterial species is listed in Table 6. In another study, ZrO_2 -ZnO nanoparticles were prepared by the sol-gel method and further evaluated for their antibacterial activity.¹⁰⁸ The mechanism of antibacterial activity by ZrO_2 NPs was proposed by Asha *et al.*⁴³ against *Pseudomonas aeruginosa* due to the rapid attraction between the negatively charged cell walls of *Pseudomonas aeruginosa* and positively charged zirconia NPs. Khan *et al.*¹⁰⁹ prepared glutamic acid-functionalized ZrO_2 NPs by using a solvothermal method and studied their antimicrobial activity against four strains of oral bacteria, *viz.*, *Streptococcus mutans*, *Rothia dentocariosa*, *Streptococcus mitis*, and *Rothia mucilaginosa*. The alteration of the surface of ZrO_2 with glutamic acid increases the dispersion characteristics of nanoparticles in an aqueous medium which in turn possibly increases interactions between bacteria and nanoparticles.

Derbalah *et al.*¹¹⁰ reported antifungal activity against *Rhizoctonia solani* which is known to be responsible for root rot in cucumber. The induction of resistance was examined using a real-time polymerase chain reaction. A concentration of 100 $\mu\text{g L}^{-1}$ of zirconia nanoparticles showed better inhibition to improve the growth of the cucumber plant than the untreated cucumber plant. Ahmed *et al.*⁵³ reported the synthesis of biogenic ZrO_2 NPs by using *Enterobacter* sp. and studied their antifungal activity against *P. versicolor* (Fig. 9). In this study, nanoparticles ruptured the cells of pathogens by getting absorbed on the cell membrane and demonstrated a significant inhibition zone at the concentration level of 20 $\mu\text{g mL}^{-1}$. Similarly, Balaji *et al.*⁶⁸ demonstrated the synthesis of spherical zirconia nanoparticles (~ 9 to 11 nm) by using the leaf extract of *Eucalyptus globulus*. In this study, the antioxidant activity of the prepared material was also investigated using human colon carcinoma (HCT-116) and human lung carcinoma (A-549) cell lines by DPPH assay. The antioxidant studies of the prepared

zirconia NPs displayed significant inhibition (85.6%) at the concentration level of 158.15 $\mu\text{g mL}^{-1}$. The cytotoxicity of the prepared ZrO_2 was also investigated against HCT-116 colon cancer and A-549 lung cancer cell lines. A mechanism for the action of ZrO_2 towards cytotoxicity is also given in this report where ROS have been identified in key roles (Fig. 10). Furthermore, the anticancer activity of zirconia-based nanomaterials against various cell lines, and their % reduction in cell viability are summarized in Table 7.

Based upon data listed in Tables 3 and 5, it can't be concretely concluded that a particular phase of zirconia would be highly efficient for antibacterial and anticancer activity as most of the studies have been carried out on different microorganisms and cell lines by different methods. However, *prima facie*, the tetragonal phase seems to be more efficient as it had shown antibacterial activity on a range of microorganisms and significant anticancer activity on a range of cell lines. But, composites of zirconia with silver or yttria or iron oxide have shown promising results for both antibacterial as well as anticancer activities. Probably, composites of the tetragonal phase of zirconia could show enhanced superior activities.

1.2.3 Adsorption. Zirconia-based nanomaterials offer a significant prospect for efficient adsorption of toxins (dyes/organic and inorganic pollutants) in wastewater treatment because they could be designed with high specific surface area, sorption active sites, small particle size, and porous structure as compared to their bulk counterparts. Such materials have been employed as adsorbents for the adsorption of various types of pollutants including dyes such as methylene blue *etc.*, heavy metal ions such as arsenic, phosphates, vanadium ions, and pharmaceutical intermediates and products from aqueous solutions. The key parameters and optimized reaction conditions from recent reports on zirconia-based adsorbents have been summarized in Table 8. Debnath *et al.*⁶⁶ demonstrated the removal of tetracycline from wastewater by using zirconia nanoparticles under an optimized pH of 6.0 and an equilibrium time of 15 min as per the scheme given in Fig. 11. In this study, the mechanism of adsorption involves strong electrostatic interaction between the protonated surface of ZrO_2 NPs and the zwitterion of tetracycline. The adsorption capacity of ZrO_2 NPs was found to be 526.32 mg g^{-1} . Ghosh *et al.*¹¹⁵ synthesized





Table 6 Antimicrobial activity of zirconia-based nanomaterials

| Zirconia-based nanomaterial, phase, shape, average size (nm), size range (nm) | Method | Microorganism | % Activity | Inhibition zone | Applications | Ref. |
|--|---|---|--|--|---|------|
| ZrO ₂ , tetragonal, spherical, -, 50–100 | Agar diffusion method | <i>S. aureus</i> , <i>E. coli</i> (bacteria) and <i>C. albicans</i> , <i>A. niger</i> (fungus) | — | 23, 32, 22, 18 mm respectively | — | 67 |
| Fe ₃ O ₄ -doped ZrO ₂ , tetragonal, spherical, 30 | Agar well diffusion method | <i>Bacillus subtilis</i> (bacteria) | — | 32 mm | Dental fillers | 104 |
| ZrO ₂ , -, spherical, 5, - | Agar well diffusion method | <i>Bacillus subtilis</i> , <i>E. coli</i> , and <i>Salmonella typhi</i> (bacteria) | — | 21, 19, 19 mm respectively | — | 105 |
| Polyaniline-ZrO ₂ , -, -, -, - | Agar diffusion method | <i>E. coli</i> and <i>Staphylococcus aureus</i> (bacteria) | — | 14, 18 mm respectively | Used in the treatment of wastewater effluent and decontamination of rearing water | 106 |
| ZrO ₂ , monoclinic, spherical, 35, 32–38 | Disc diffusion method | <i>Pseudomonas aeruginosa</i> , <i>E. coli</i> , <i>Staphylococcus aureus</i> , and <i>Bacillus subtilis</i> (bacteria) | — | 10 mm in <i>Pseudomonas aeruginosa</i> | — | 43 |
| ZrO ₂ -ZnO, -, spherical, 26–34, - | Disc diffusion method | <i>Bacillus subtilis</i> , <i>Streptococcus mutans</i> , <i>Staphylococcus aureus</i> , <i>E. coli</i> , <i>Pseudomonas aeruginosa</i> , and <i>Klebsiella oxytoca</i> (bacteria) | — | — | — | 108 |
| Glutamic acid functionalized zirconia (GA-ZrO ₂), cubic, spherical, 2.5, - | Autoclaved brain heart infusion broth (BHI) or agar | <i>Rothia mucilaginosa</i> , <i>Rothia dentocariosa</i> , <i>Streptococcus mitis</i> , and <i>Streptococcus mutans</i> (bacteria) | 58% ± 4.9% (<i>R. dentocariosa</i>) 52% ± 2.9% (<i>Streptococcus mutans</i>) | — | — | 109 |
| ZrO ₂ , -, spherical, -, 15–21 | Disc diffusion method | <i>Bacillus subtilis</i> , <i>Staphylococcus aureus</i> , <i>E. coli</i> and <i>Pseudomonas aeruginosa</i> (bacteria) | — | 20 mm in <i>Pseudomonas aeruginosa</i> | Anti-tooth decay applications | 107 |
| ZrO ₂ , tetragonal, -, 17, - | Agar dilution method | <i>Rhizoctonia solani</i> (fungus) | 86.6%, 100 µg L ⁻¹ | — | Improved cucumber growth | 110 |
| ZrO ₂ , -, spherical, -, 33–75 | Agar well diffusion | <i>Pestalotiopsis versicolor</i> (fungus) | — | 25.18 ± 1.52 mm | Protects bayberry (twig blight disease pathogen) plants against antifungal agents | 53 |

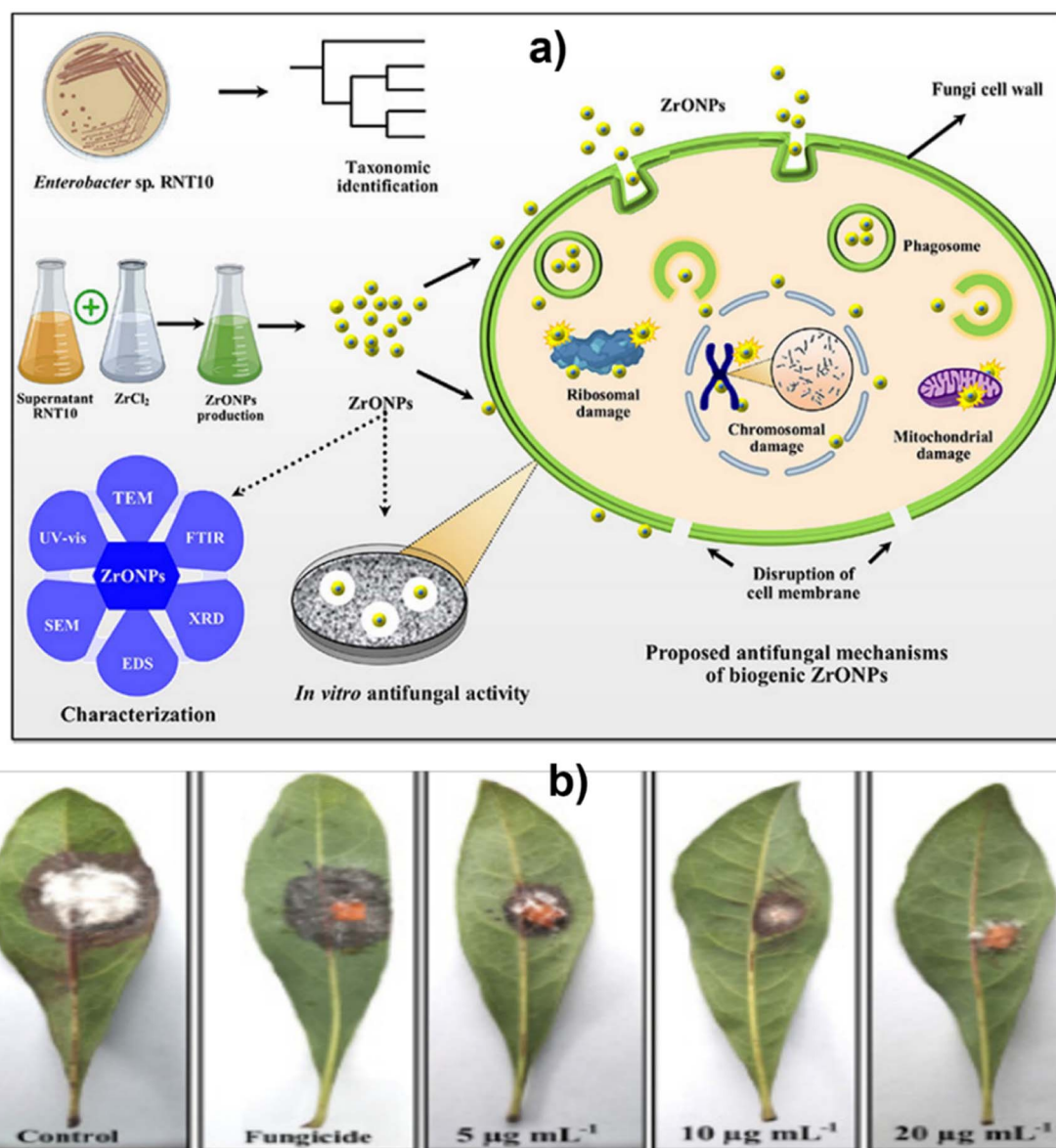


Fig. 9 (a) Green synthesis and characterization of zirconium oxide nanoparticles by using a native *Enterobacter* sp. and its antifungal activity against bayberry twig blight disease pathogen *Pestalotiopsis versicolor* and (b) effect of various concentrations of ZrO NPs on bayberry leaves infected with *P. versicolor* strain XJ27. Adapted from ref. 53 with permission from Elsevier, copyright 2020.

a ceria-incorporated ZrO_2 nanocomposite having a particle size of ~ 50 nm and employed it for the removal of arsenic(III) from the aqueous phase. The prepared material offers its surface for the oxidation of arsenic(III) to arsenic(V). It showed a kinetically pseudo-second-order reaction and the obtained Langmuir isotherm indicated a monolayer adsorption capacity of 17.07 mg g^{-1} .

It can be suggested from the literature summarized in Table 8 that both bare ZrO_2 and its composites could be employed for the adsorption of pollutants from effluents effectively. Specifically, designing composites of ZrO_2 NPs with graphene oxide could be a useful approach for fabricating economical and efficient adsorbents for a range of pollutants and industrial-by-products.

1.2.4 Dentistry. Nanoscale ZrO_2 has increasingly become one of the significant ceramic materials in the field of dentistry due to its outstanding properties such as natural color, stable chemical properties, good corrosion resistance, high toughness, mechanical strength and, most importantly, its superior biocompatibility. It is being used as a dental implant biomaterial and dental crown. Zirconia-based ceramics are found to be chemically inert materials, therefore, considered preferable candidates for superb cell adhesion, and no unfavourable systemic reactions have been observed. Modification of the implant surface is highly necessary to increase its strength and prevent its failure caused due to insufficient bone formation and apposition. Methods such as sandblasting, acid etching, and laser etching have been performed for this purpose.¹²²



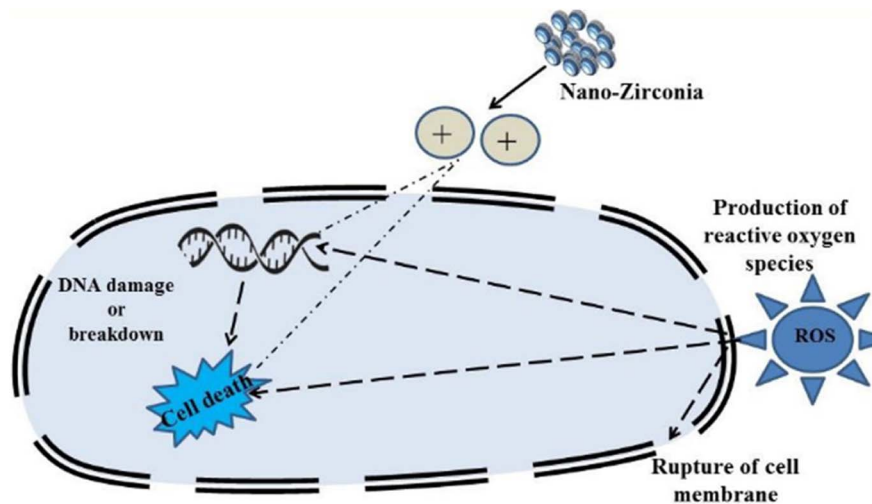


Fig. 10 Plausible mechanism of action by ZrO_2 NPs towards cytotoxicity. Adapted from ref. 68 with permission from Elsevier, copyright 2017.

Table 7 Anticancer activity of zirconia-based nanomaterials

| Zirconia-based nanomaterial, phase, shape, average size (nm), size range (nm) | Method | Cell line | % Reduction in cell viability | Ref. |
|---|--------------|---|--|------|
| ZrO_2 , tetragonal, spherical, -, 9–11 | MTT | Human colon carcinoma (HCT-116) and human lung carcinoma (A-549) cell lines | 50% | 68 |
| ZrO_2 , tetragonal, tetragonal, 56.8, - | MTT | MCF-7 cell lines | 18% | 65 |
| $Fe_3O_4@ZrO_2$, -, spherical, 155.23 ± 7.1 , - | CCK-8 method | HepG2 cell line | $54.76\% \pm 5.88\%$ to $17.72\% \pm 3.66\%$ | 111 |
| AgNPs@UiO-66(DMF), -, cubical, 140, - | MTT | SMMC-7721 and HeLa cells | 92 and 89% respectively | 112 |
| Y_2O_3/ZrO_2 , -, -, 33.9 ± 2.1 , - | MTS and NRU | Human skin keratinocyte (HaCaT) cells | 94%, 90%, 78% and 65% (in 24 h) | 113 |
| Sulphated zirconia, -, tetragonal, 43, 37–54 | MTT | HT29, human breast cancer MCF-7, and human liver cancer HepG2 | 50% | 114 |

Aivazi *et al.*¹²³ reported the laser method for the modification of the surface to support osseointegration. In this study, surface modification caused a positive impact on the viability of fibroblast L929 cells. Tahsili *et al.*¹²⁴ prepared meso-/micro-/nanoscale rough zirconia by solid-state laser sculpting for improving the performance of zirconia implants. The prepared zirconia shows a significant increase in osseointegration. Yu *et al.*¹²⁵ studied the effect of surface properties of nano-scale zirconia on the adhesion and biofilm development of *Streptococcus mutans*. In this study, it was observed that the surface roughness of the nanostructure can influence the adhesion force and early attachment (2 and 4 h) of *Streptococcus mutans* on the zirconia. Nevertheless, it did not affect the formation of biofilm in the later stages of 6, 8, 12, and 24 h.

Recently, an yttria-stabilized t-zirconia polycrystal material has been used for implants and dental applications due to its excellent mechanical properties.⁷ Interestingly, zirconia-based nanomaterials can also be used to improve the physical and mechanical properties of polymethyl methacrylate (PMMA) which is a denture repair material. Gad *et al.*¹²⁶ also demonstrated that the tensile strength of PMMA could be increased

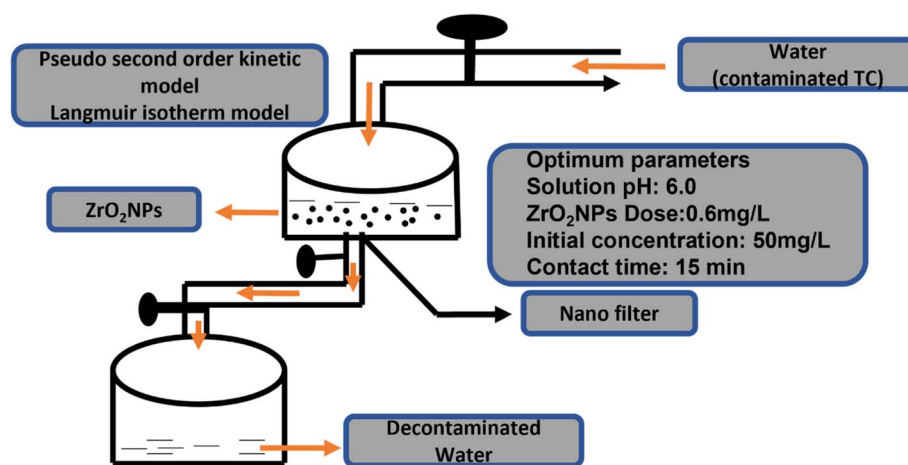
using zirconia nanoparticles and can be used as a denture base with reduced translucency. Lohbauer *et al.*¹²⁷ demonstrated the preparation of spherical ZrO_2 NPs in the range of 20–50 nm using laser vaporization. These nanomaterials were used as fillers in the dental adhesive.

Fathima *et al.*¹⁰⁷ synthesized zirconia nanoparticles and demonstrated their excellent antibacterial activity and anti-decay application on teeth. Bacteria may generate acids during their metabolic activities which result in the demineralization of teeth (Fig. 12). However, coating with zirconia NPs has shown resistance to acid attacks as well as remarkable activity against oral pathogens bearing negatively charged cell walls. The electrostatic interactions between positively charged nanoparticles and charged microorganisms inhibit the activities of microbes. Gad *et al.*¹²⁸ reported a novel method for denture stomatitis inhibition by using ZrO_2 nanoparticles owing to their antifungal activity. The authors investigated the inhibitory effect of ZrO_2 nanoparticles on *Candida albicans* adhesion to restore polymethyl methacrylate denture bases. Similarly, in another investigation by Oshima *et al.*,¹²⁹ a zirconia/alumina composite nanomaterial (Ce-TZP/ Al_2O_3) was found to increase the



Table 8 Zirconia-based nanomaterials as adsorbents

| Adsorption activity | Zirconia nanomaterial type | Experimental conditions (temperature, pH, and equilibrium time) | Adsorption capacity | Isotherm | Ref. |
|--|--|---|---|-------------------------------------|------|
| Tetracycline from wastewater | ZrO ₂ | –, 6, 15 min | 526.32 mg g ^{–1} | Langmuir isotherm | 66 |
| Phosphate in treated wastewater | PANI-ZrO ₂ | –, 25 °C, 6.9, – | 32.4 mg g ^{–1} | Langmuir isotherm | 106 |
| Phospholipid adsorption from jatropha oil used for biofuel | ZrO ₂ | – | 13.92 mg g ^{–1} | Langmuir isotherm | 77 |
| Adsorption of methylene blue and tetracycline hydrochloride from water | UiO-66-(OH) ₂ /graphene oxide (GO) | 298 K, 9, – | 99.96% and 94.88%, respectively | Langmuir, Freundlich, Sips isotherm | 116 |
| Adsorption of arsenic(III) from the water | Ceria incorporated zirconia nanocomposites | 303 K, 7.0, – | 17.07 mg g ^{–1} | Langmuir isotherm | 115 |
| Adsorption of chromium and dyes RB-21, RR-141, and Rh-6 G | Chitosan-zirconium phosphate nanostructures | 303 K, –, – | 311.53 mg g ^{–1} , 457 mg g ^{–1} , 435.1 mg g ^{–1} and 438.596 mg g ^{–1} respectively | Langmuir isotherm | 117 |
| Adsorption of vanadium ion | CNZ@NCT/Zr | 298 K, <5, – | 277.75 mg g ^{–1} | Langmuir isotherm | 118 |
| Adsorption of phosphate from water | ZrO ₂ /Fe ₃ O ₄ | –, neutral, – | 29.5 mg g ^{–1} | Langmuir isotherm | 119 |
| Adsorption of norfloxacin in water | UiO-66-NH ₂ | 298 K, 6–8.5, 6 h | 222.5 mg g ^{–1} | Langmuir isotherm | 120 |
| Adsorption of toxic arsenite (As(III)) and arsenate (As(V)) from aqueous solutions | Mesoporous zirconia nanostructure | –, neutral, – | 105.03 and 110.29 mg g ^{–1} , respectively | Langmuir isotherm | 121 |

Fig. 11 Schematic diagram of batch adsorption of TC on the surface of biosynthesized ZrO₂ NPs.⁶⁶

osseointegration capability and osteogenic response of implants, where the surface of the nanostructured Ce-TZP/Al₂O₃ was modified by HF treatment.

Yamada *et al.*¹³⁰ prepared silver NP-coated yttria-stabilized ZrO₂ and investigated its antibacterial activity against oral

pathogens such as *Staphylococcus aureus*, *Streptococcus mutans*, *Escherichia coli*, and *Aggregatibacter actinomycetemcomitans* (Fig. 13). The composite shows antimicrobial activity, particularly against *Escherichia coli*. The efficacy of the prepared material was tested *via* an accelerated aging test. In this test, the



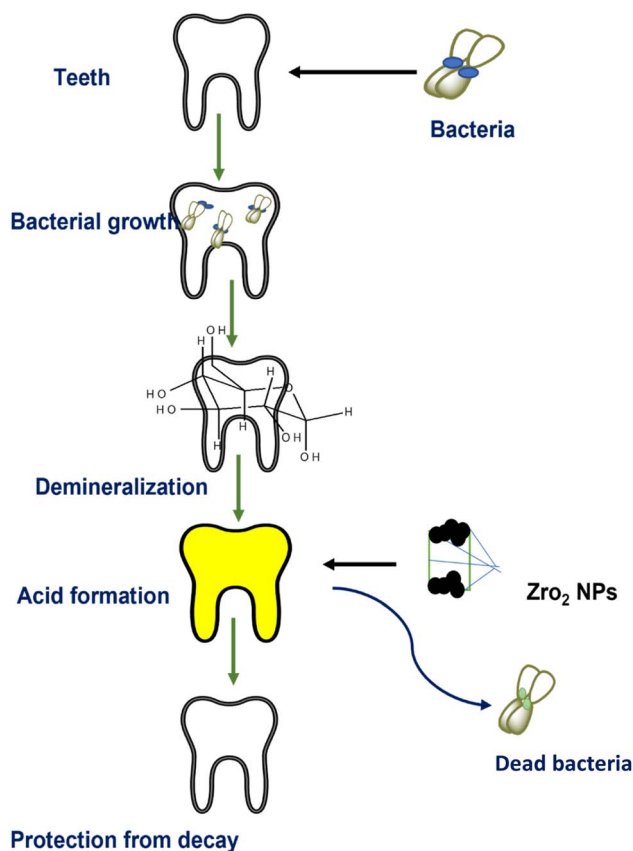


Fig. 12 Role of ZrO₂ NPs in dental care.¹⁰⁷

release profile of silver NPs was investigated which was found to be less than 1 percent, hence the prepared material was found to be thermally stable. Therefore, it was suggested that the bactericidal action of the composite was due to its contact with the bacteria and not owing to the release of silver NPs. Thus, the prepared material could be further investigated for the treatment of dental caries and periodontal disease.

1.2.5 Sensing. Zirconia-based nanomaterials are also being explored for sensing applications due to their favorable electrochemical and physicochemical properties such as high surface area, chemical stability, and fluorescent nature. Various sensing parameters, modes of generating signals and

substances for which such materials have been employed have been summarized in Table 9. For instance, Zhang *et al.*¹³¹ demonstrated the detection of trace amounts of ethanol using zirconia nanoparticles with high sensitivity. The developed sensor indicated a constant and stable response for 100 h *via* a chemiluminescence reaction and demonstrated remarkable durability. The formation of acetaldehyde as one of the possible luminescent intermediates was observed in the sensing process.

Garadkar *et al.*¹³² fabricated thick film resistors using nanostructured ZrO₂ (25 nm) *via* the screen printing technique. The as-prepared resistor was used for the selective detection of H₂ gas and showed a significant signal at room temperature even in the presence of a smaller amount of H₂. It was reported that films exhibited a clear response to 50 ppm of H₂ in contrast to other gases at a higher amount of 1000 ppm. Chikere *et al.*¹³³ fabricated a zirconia nanoparticle modified carbon paste electrode and employed it for the electrochemical detection of gallic acid in a concentration range of 1×10^{-6} to 1×10^{-3} mol L⁻¹. The weak ionic interactions between the oxygen anions of zirconia nanoparticles (monoclinic) and positively charged graphite sheets, as depicted in Fig. 14, were found to stabilize the sensor. The designed sensor generated a signal *via* electrochemical oxidation of gallic acid and was employed for detection of the latter in red and white wine.

Wang *et al.*¹³⁴ prepared zirconia nanoparticles having an average size of 20 nm and employed them to make a thick film humidity sensor using a substrate. The prepared sensor showed a change in impedance from 10^6 to 10^2 Ω with an increase in relative humidity (RH) from 11 to 98%. In this report, the impact of temperature on the impedance of the sensor was evaluated and the conduction process of the sensor was explained using an equivalent circuit with resistors and capacitors. Gong *et al.*¹³⁵ fabricated a nanostructured composite of zirconia and graphene by solid-phase extraction. The prepared material afforded high recognition and enrichment for phosphoric moieties due to the NPs of ZrO₂, large surface area, and high conductivity owing to graphene nanosheets. The prepared material was successfully employed for the detection of methyl parathion *via* square-wave voltammetry.

Trinadh *et al.*¹³⁶ prepared an electrochemical sensor using graphene quantum dots (GQDs) and lanthanum doped zirconia

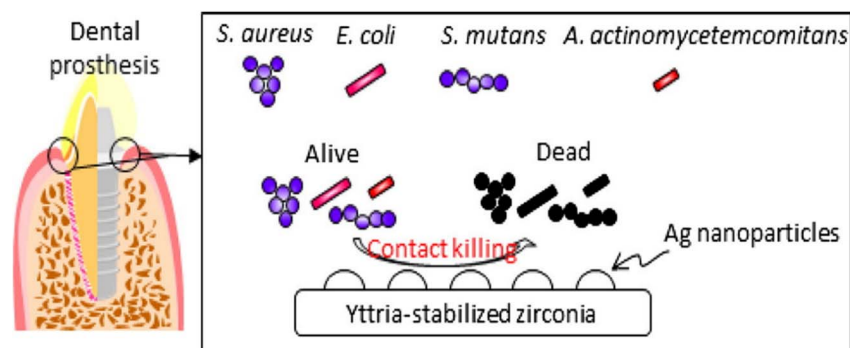


Fig. 13 Silver NP-coated zirconia for antibacterial prosthesis. Adapted from ref. 130 with permission from Elsevier, copyright 2017.



Table 9 Application of zirconia-based nanomaterials in sensing

| Sensor | Synthesis method | Analyte | Working temperature | Recovery time, response time | Sensing technique | Limit of detection/linear response | Ref. |
|--|---------------------------------------|------------------|---------------------|------------------------------|---|---|------|
| ZrO ₂ on a silicon substrate | — | Humidity | 10–30 °C | 60 s, 130 s | Measurement of impedance | — | 134 |
| ZrO ₂ NP decorated graphene nanosheets | Electrochemical approach | Methyl parathion | — | — | Square-wave voltammetry | 0.6 ng mL ⁻¹ | 135 |
| Zirconia nanocubes | Hydrothermal method | Arsenic(III) | — | —, below 2 s | Cyclic voltammetry | 5 ppb | 29 |
| Zirconia NPs | Precipitation method | Dimethyl amine | 330 °C | —, less than 50 s | Cataluminescence | 6.47 × 10 ⁻⁴ ng mL ⁻¹ | 138 |
| ZrO ₂ 5%Y/ZrO ₂ Ag-5% Y/ZrO ₂ | Hydrothermal method | CO ₂ | 300–400 °C | 22 s, — | Measurement of current | — | 139 |
| Polyaniline zirconia nanocomposite | Polymerization method | Esomeprazole | — | —, 9 s | Electrochemical impedance spectroscopy and cyclic voltammetry | 97.21 ng mL ⁻¹ | 140 |
| Au-NPs/ZrO ₂ nanofiber | Nonthermal spin-coated sol-gel method | Pesticide | — | —, — | Surface-enhanced Raman scattering | 10 ⁻⁶ to 10 ⁻⁸ M | 137 |
| GQDs@La ³⁺ @ZrO ₂ nanoparticles | Bottom-up and sintering approach | Flutamide | Room temperature | — | Electrochemical signal | 0.00082 μM | 136 |
| ZrO ₂ NP modified carbon paste electrode | Hydrothermal method | Gallic acid | Room temperature | — | Differential pulse voltammetry (DPV) | 1.24 × 10 ⁻⁷ mol L ⁻¹ | 133 |
| SiO ₂ /ZrO ₂ composite | Sol-gel method | NO ₂ | 25 °C | Varied | Measurement of resistance | — | 141 |

nanoparticles for the detection of flutamide in urine samples as per the scheme depicted in Fig. 15. For designing this sensor, La³⁺@ZrO₂ was prepared by the coprecipitation method using a solution of zirconyl chloride and lanthanum nitrate as metal precursors and NH₄OH as the precipitating agent. The addition of NH₄OH decreases the pH of the solution which leads to the formation of precipitates. The obtained precipitates were washed with methanol and water, dried, and calcined at 600 °C. Finally, the glass carbon electrode-based sensor was fabricated by drop casting a suspension of 1.0 mg of GQDs and 1.0 mg of La³⁺@ZrO₂ on a pre-treated GCE. In this study, the signal was generated due to the electrochemical reduction of flutamide (12 μM) in phosphate buffer solution (pH = 7).

Lee *et al.*¹³⁷ demonstrated the preparation of Au nanoparticle coated zirconia nanofibers and employed them as a surface-enhanced Raman scattering sensor for the detection of a variety of pesticide (carbaryl, phosmet, cypermethrin, and permethrin) residues in fruits or vegetables. In this study, nanofibers of zirconia were fabricated on a silicon wafer *via* the spin coating technique using a solution of ZrCl₄ in isopropanol. Afterward, gold NPs were deposited on ZrO₂ nanofibers *via* an electron beam evaporator. The prepared sensor shows excellent selectivity to organophosphates even in the presence of multiple pesticides. It was successfully employed for a practical application using diluted juice containing sliced pesticide-containing apple peels and could differentiate different pesticides.

The literature summarized in Table 9 revealed that ZrO₂ or its composites could be employed for designing a range of sensors based upon different techniques such as electrochemical, luminescence *etc.* However, significant LODs have been displayed by them *via* electrochemical signals. Specifically, a cyclic voltammetry-based signal from such sensors seems to be a more promising one owing to the electrical properties of ZrO₂.

2 Outlook and futuristic scope

Recently, there have been substantial efforts and significant developments in the field of synthesis of nanomaterials for better exploitation in numerous applications. With the aid of a modified outlook of nanomaterials and the comprehension of their potential in various industries, including, but not limited to, pharmaceuticals, textiles, paper, pulp, and drug delivery, the influx of zirconia nanoparticles has offered innumerable benefits and provided solutions for several continuing problems related to clinical diagnosis, catalysis, and dentistry, and can indefinitely offer long-term economic benefits. In the particular interest towards the synthesis of zirconia nanoparticles, several chemical and green methods have been developed and their plausible mechanisms have been studied in detail. Furthermore, efforts have also been made to achieve monodispersity, desired particle size, shape, and stability *via* tuning the reaction conditions and parameters. However, there are challenges associated with improving the efficiency of the synthesis and achieving the specific particle sizes and morphology, a desirable cubic phase, long-term stability, and reduced aggregation of particles. Moreover, expansion of the volume of particles with the cubic phase has also been observed



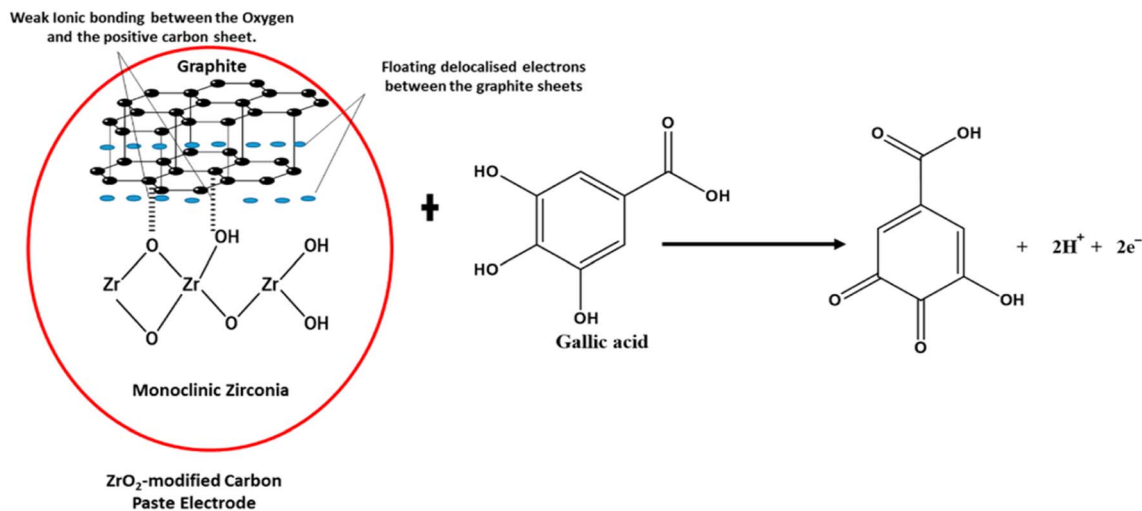


Fig. 14 Proposed interaction between ZrO₂ NPs and graphite, demonstrating a weak ionic bond between some of the oxygen and the positive graphite sheet in the electrochemical determination of gallic acid.¹³³

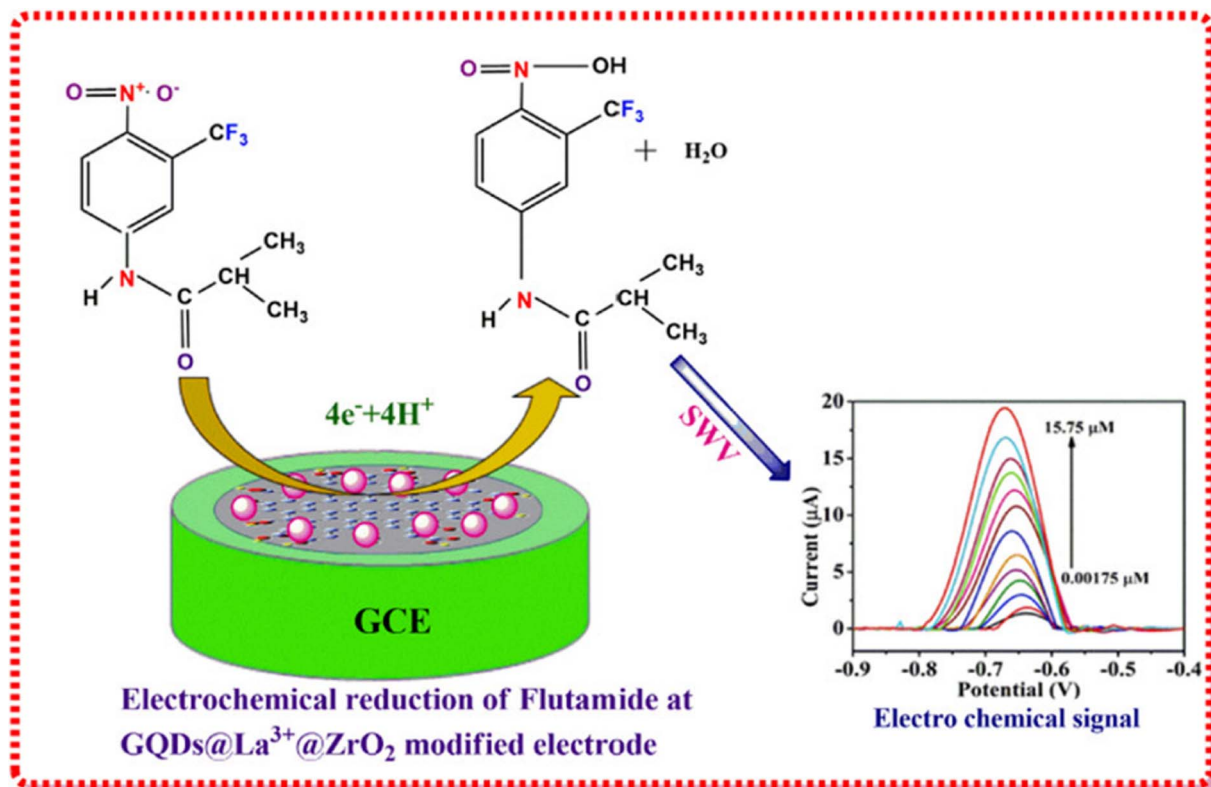


Fig. 15 Pictorial representation of the fabrication of GQDs@La³⁺@ZrO₂@GCE. Adapted from ref. 136 with permission from Elsevier, copyright 2017.

when employed in high-temperature applications due to its transformation into the tetragonal phase and subsequently tetragonal to monoclinic phase which results in the generation of stress and cracks and makes the material unfit for applications. Therefore, a long journey is still pending in this direction to overcome the above-mentioned limitations. To gain an insight, a significant number of theoretical studies using density functional theory have been performed to understand

the structure, nature of surface sites, and stability of zirconia nanoparticles. Recently, Maleki *et al.*¹⁴² performed DFT calculations to obtain the ¹⁷O solid-state NMR chemical shifts for studying the nature of surface sites on zirconia. In an interesting study, Puigdollers *et al.*¹⁴³ showed that nanoparticles with partly truncated octahedral morphology display the highest stability. Moreover, nanoparticles with only stable facets {101} were found to be less stable due to the presence of reduced Zr,



thereby confirming that low-coordinated sites create defective states in the electronic structure and reduce the effective bandgap, which can result in enhanced interaction with deposited species and could modify the photocatalytic activity. In another report, the presence of acidic and basic sites on zirconia was investigated using DFT calculations for obtaining the IR and NMR spectra of adsorbed molecules (carbon monoxide and pyrrole). This study confirms the presence of acidic Zr sites and basic O⁻ sites in nanoparticles. However, further theoretical studies for obtaining stronger acidic and basic sites and achieving the stability of the cubic phase in zirconia nanoparticles could enhance the suitability of zirconia in various applications as discussed earlier. To ensure the commercial applicability of such materials in various technological applications, future perspectives should focus not only on an in-depth understanding of the surface structure and active sites but also on the design and development of well-defined chemical and green systems for the substantially controllable site environment and attractive characteristics.

Experimentally, two approaches have been employed to enhance the stability and applications of zirconia nanoparticles. The first one involves *in situ* modification with stabilizers and other metal oxides such as CaO, MgO, and Y₂O₃ during the preparation of the nanomaterials. Modification with another metal oxide could generate vacancies stimulating changes in the electronic and geometric environment that can further enhance the optical and electrical properties which is also supported by several theoretical studies.^{143–145} The second strategy involves exploring greener methods with novel additives that could be employed. Although the synthesis of zirconia nanomaterials using plant extracts and microbes is considered to be clean, nontoxic, simple, and safe, there are only a few reports in this regard and sufficient efforts have not been made with diverse varieties of plants and microorganisms with different possible reaction pathways. Moreover, the combination of the sol–gel method and greener reactants is rarely reported. Specifically, the sol–gel method offers unique advantages such as good control over the structure and kinetics of the process. Moreover, fine control of the composition of a multi-component nanomaterial could be achieved *via* this method as small quantities of dopants can be introduced in the sol and uniformly dispersed in the final product. Therefore, further improvising this method using greener ingredients by keeping in view the motive of increasing the strength and biocompatibility could result in the invention of new materials suitable for various applications such as dentistry, catalysis, sensors, *etc.*

Conflicts of interest

There are no conflicts to declare.

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