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Graphdiyne oxide doped SnO₂ electron transport layer for high performance perovskite solar cells†

 Lili Yao,^{‡a} Min Zhao,^{‡ab} Le Liu,^{ab} Siqi Chen,^b Jin Wang,^b Chengjie Zhao,^{ab}
 Zhiyu Jia,^{id *c} Shuping Pang,^{id a} Xin Guo^{id d} and Tonggang Jiu^{id *ab}

High-performance planar perovskite solar cells (PSCs) are dependent on the properties of electron transport layers (ETLs). Here, we report an effective interface engineering strategy by incorporating an oxidized form of graphdiyne into a SnO₂ ETL. In addition to the unique structural properties of graphdiyne, graphdiyne oxide (GDYO) is characteristic of additional hydrophilic carboxy and hydroxy functional groups that may form chemical bonds with uncoordinated Sn. The interaction thereof is able to reduce the oxygen vacancy within SnO₂, thereby passivating the surface defects of SnO₂; this promotes electron transport while also suppressing non-radiative recombination. Notably, the work function of the GDYO-doped SnO₂ film matches well with the perovskite conduction band, resulting in a high open circuit voltage. As a consequence, the GDYO-containing device demonstrated a high PCE of 21.23% with a V_{OC} of 1.13 V, a J_{SC} of 24.49 mA cm⁻² and an FF of 76.85%, superior to those of the control device without GDYO. Furthermore, the unencapsulated device maintained 84% of the initial efficiency after 80 °C for 24 days and 71% after continuous illumination for 160 h. This work provides guidance for developing efficient and stable perovskite devices from the perspective of optimizing interface properties, presenting great potential of functionalized graphdiyne for practical applications.

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1. Introduction

Organic–inorganic hybrid perovskite solar cells (PSCs) have been widely investigated because of multiple advantages of the perovskite absorption layer itself, such as excellent charge-carrier mobility, broad range for light absorption with a high extinction coefficient, ambipolar transport properties, and long electron–hole diffusion length.^{1,2} Over the past decade, great breakthroughs have been achieved in the field of PSCs. Not only the photoelectric conversion efficiency (PCE) has been raised from 3.8% to 25.5%, the stability of PSCs has also been improved.^{3–5} These achievements propel PSCs a big step toward practical applications. Generally, the ambipolar

semiconducting characteristics of perovskite materials endow PSC devices with conventional mesoscopic and planar structures, such as n-i-p and p-i-n configurations.^{6,7} In comparison with the latter, the n-i-p PSCs have been more extensively studied, and the main factors that determine their performance are the optical, electrical and structural properties of their electron transport layers (ETLs).^{8,9}

Metal oxides are considered to be excellent ETLs because of their good photoelectric properties, high stability and low cost.¹⁰ Among various metal oxides used as electron-transport layers, TiO₂ is widely studied for its low cost and favorable charge transfer ability.^{11–13} However, removing the residual organic components in a TiO₂ ETL during high temperature sintering will enhance the crystallinity of TiO₂, which obviously complicates the fabrication process and confines its photovoltaic applications.¹⁴ To avoid this shortcoming of TiO₂, an alternative ETL of SnO₂ has been paid particular attention owing to its compatible energy level with perovskite, high electron mobility of 249 cm² V⁻¹ s⁻¹, wide optical bandgap (3.6–4.0 eV), high transmittance over the entire visible spectrum, and excellent stability under a variety of test conditions.¹⁵ Moreover, it is easier to prepare SnO₂ by a solution process at low temperature, which is compatible with the manufacturing techniques of flexible solar cells, tandem solar cells, and large-scale commercialization.¹⁶ Although SnO₂ shows favorable application potential in inverted PSC devices, some flaws of

^a Qingdao Institute of Bioenergy and Bioprocess Technology, Chinese Academy of Sciences, Qingdao, 266101, China. E-mail: tgjiu@sdu.edu.cn

^b Science Center for Material Creation and Energy Conversion, Institute of Frontier and Interdisciplinary Science, Shandong University, Qingdao, 266237, China

^c MOE Key Laboratory of Cluster Science, School of Chemistry and Chemical Engineering, Beijing Institute of Technology, Beijing, 100081, China. E-mail: jzy@bit.edu.cn

^d State Key Laboratory of Catalysis, Dalian Institute of Chemical Physics, Chinese Academy of Sciences, Dalian National Laboratory for Clean Energy, Dalian, 116023, China

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‡ These authors contributed equally to this work.

using metal oxides as ETLs are still present, such as an incontinuous interface with pinholes and rough surface and oxygen vacancy defects caused by the presence of uncoordinated Sn.¹⁷ These interface defects are likely to act as centers of charge recombination, consequently leading to poor performance of SnO₂-based devices.¹⁸ Fortunately, numerous studies have shown that interface defects can be modified by doping carbon nanomaterials, such as graphene,¹⁹ fullerene,²⁰ and carbon nanotubes.²¹

Graphdiyne (GDY) is a new member of the carbon nanomaterials family which consists of sp- and sp²-hybridized carbon and possesses a highly delocalized π system.^{22–24} These characteristic electronic structures endow GDY with excellent semi-conducting properties, superior electrical properties, high stability, *etc.*; thus, GDY is widely used in various fields,

including energy storage and conversion and biochemical sensing.^{25–33} To date, a number of studies about the application of GDY in PSC devices have been reported. For instance, doping GDY into PCBSD can improve the orientation of perovskite films and further enhance their growth and crystallization.³⁴ GDY was also used to modify PC₆₁BM, resulting in improved conductivity and interfacial coverage of an ETL.³⁵ In addition, GDY-modified P3CT-K was introduced as a hole transport layer (HTL) for improved device performance, with enhanced PCE from 16.8% to 19.5%.³⁶ All the above results demonstrate that GDY doping is an effective approach to device performance optimization. Graphdiyne oxide (GDYO) as a new graphdiyne derivative retains the attractive structural characteristics of GDY; however, it is equipped with additional hydrophilic carboxy and hydroxy functional groups, enabling it

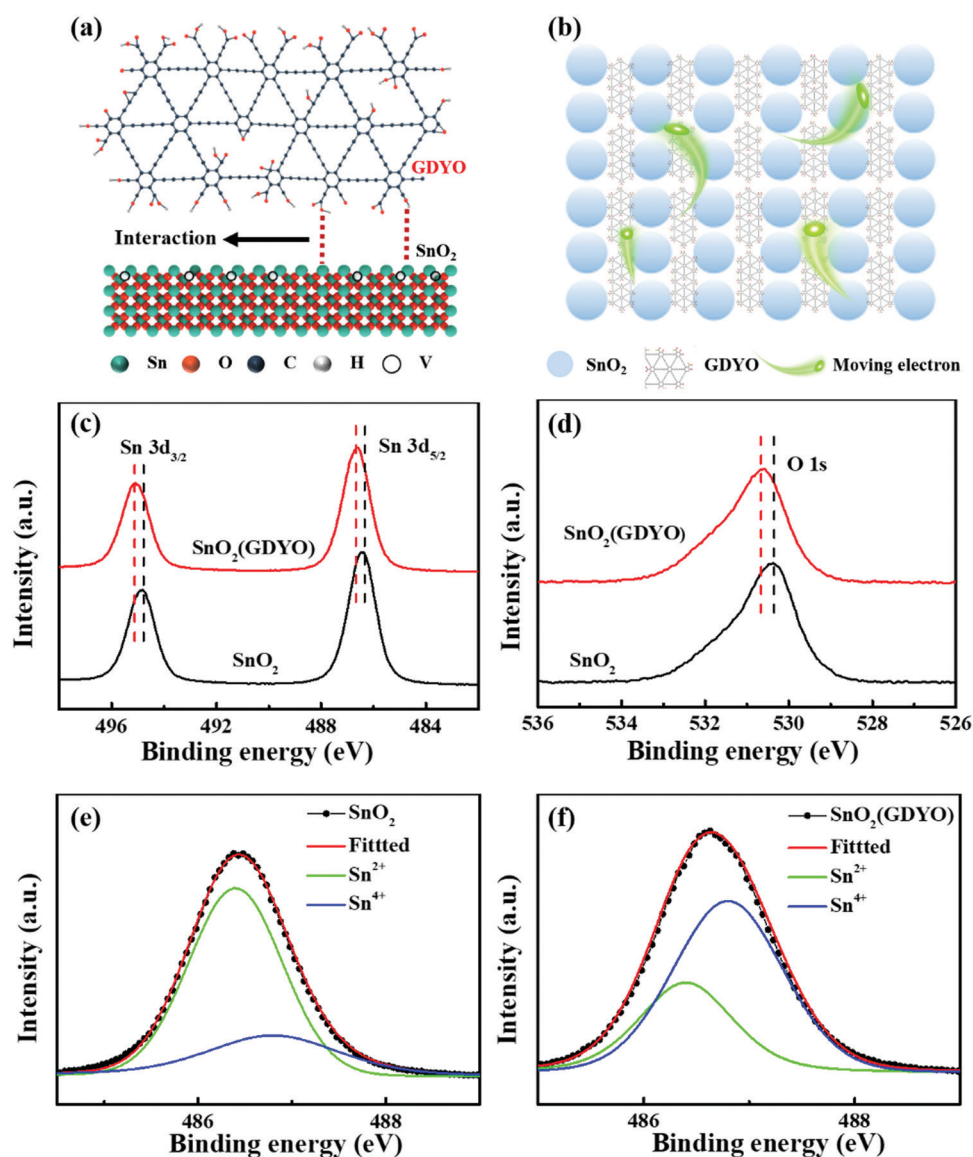


Fig. 1 Schematic of (a) the proposed interaction between the SnO₂ surface and GDYO and (b) the effect of GDYO doping in SnO₂. The XPS spectra of (c) Sn 3d, (d) O 1s; XPS spectra of the Sn 3d_{5/2} peak measured from (e) the SnO₂ layer and (f) the SnO₂(GDYO) composite layer.

to complex with other metal oxides. Compared to GDY, GDYO possesses more features that make it promising in PSC applications.³⁷

In this work, we first report a GDYO-doped SnO₂ composite material as an ETL to prepare high-performance n-i-p perovskite devices. Compared with the control device based on a (FAPbI₃)_{1-x}(MAPbBr₃)_x active layer (19.19%), the GDYO-containing device delivered an increased PCE of 21.23% with negligible hysteresis of current density–voltage (*J*–*V*) curves. It can be found that the work function of SnO₂(GDYO) was well matched with the perovskite conduction band, which was conducive to a high *V*_{OC}. Meanwhile, the hydrophilic carboxy and hydroxy functional groups of GDYO can form chemical bonds with uncoordinated Sn and reduce the oxygen vacancy in SnO₂, thus passivating the surface defects of SnO₂, promoting electron transport and reducing non-radiative recombination. Finally, the unencapsulated PSCs based on GDYO showed good stability, maintaining 84% of the initial efficiency after 24 days at 80 °C and 71% after 160 h under continuous illumination.

2. Results and discussion

Firstly, the properties of the SnO₂ and GDYO-doped SnO₂ were characterized to confirm the presence of GDYO in SnO₂ and the associated interaction. To further investigate the interaction between GDYO and SnO₂ nanoparticles, Raman spectroscopy was conducted, and the results are shown in Fig. S1 (ESI[†]). Fig. S1b (ESI[†]) exhibits that the prominent D band and G band peaks of sp² carbon were located at ~1362 cm⁻¹ and ~1582 cm⁻¹, respectively.³⁸ In the SnO₂(GDYO) composites, the D band red shifted to a lower frequency region at ~1324 cm⁻¹, which reveals the interaction between GDYO and SnO₂ nanoparticles. According to previous reports, the redshift can be attributed to the coupling between SnO₂ nanoparticles and GDYO.³⁹ Thus, we can assume that some

chemical interaction between SnO₂ and GDYO should be present. As is known, there are surface defects in the SnO₂ ETL, which are mainly caused by oxygen vacancy defects formed by uncoordinated Sn on the surface. The existence of these defects seriously affects charge extraction and transport, and it leads to poor performance of perovskite photovoltaic devices. GDYO is oxidized and exfoliated from GDY and contains various functional groups, which are expected to inactivate the oxygen vacancy defects by bonding with the uncoordinated Sn.⁴⁰ Fig. 1a shows a schematic of the proposed interaction between the SnO₂ surface and GDYO. It depicts the chemical bonding formation between the SnO₂ surface and GDYO, where the uncoordinated Sn on the surface of SnO₂ interacts with the carboxy and hydroxy groups in GDYO through chemical bonding. Fig. 1b shows the effect of this interaction, which could promote the electron transport. To demonstrate the interaction between GDYO and SnO₂ nanoparticles, X-ray photoelectron spectroscopy (XPS) was carried out to explore the state of Sn and O elements in the SnO₂(GDYO) and SnO₂ films, respectively (Fig. S2, ESI[†]). As shown in Fig. 1c, the spectra of Sn 3d can be deconvoluted into peaks of 486.36 eV and 494.89 eV, which are in agreement with Sn 3d_{5/2} and Sn 3d_{3/2}, respectively. Moreover, the peaks of Sn 3d shift slightly to higher binding energies for the SnO₂(GDYO) film. Similar behavior of the O 1s peak for the SnO₂(GDYO) film can be observed compared with that of the SnO₂ film (Fig. 1d). The shift toward a higher binding energy indicates the presence of additional negative charges around Sn and O atoms. Furthermore, taking a closer look at Sn 3d_{5/2}, two peaks appear at 486.4 eV and 486.8 eV, representing Sn²⁺ and Sn⁴⁺, respectively (Fig. 1e and f). The ratio of Sn⁴⁺ and Sn²⁺ peaks in the SnO₂(GDYO) composite layer is larger than that in the SnO₂ layer. It indicates that the density of the electron cloud around Sn increases with more Sn and O binding, and thus the oxygen-deficient defects are effectively passivated by GDYO. Meanwhile, we used XPS Gaussian

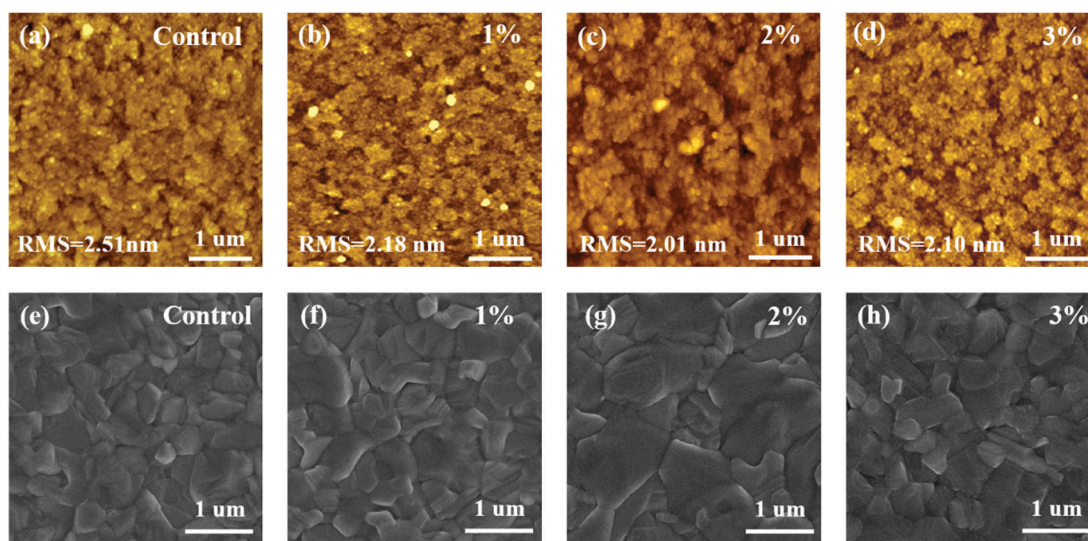


Fig. 2 AFM images of (a) SnO₂ and (b–d) SnO₂(GDYO) ETLs with different doping concentrations. SEM images of perovskite films grown on (e) SnO₂ ETL and (f–h) SnO₂(GDYO) ETLs with different doping concentrations. The scale bar is 1 μm.

distribution to conduct peak splitting treatment on O 1s to explore its valence state, vacancy and composition. According to Fig. S3 (ESI[†]), there are two oxygen states in the SnO₂ film and SnO₂(GDYO) film: 530.4 eV for the oxygen tin binding energy (O–Sn–O) and 531.4 eV for the oxygen vacancies energy. It was found that the proportion of O–Sn–O bonds was slightly increased by introducing GDYO in the SnO₂ film, and the peak area of the oxygen vacancies decreased slightly. GDYO interacts with SnO₂ to reduce the oxygen vacancy of SnO₂, thus suppressing non-radiative recombination and promoting charge transport.^{8,41}

The micro-morphology of the interface layer plays an important role in achieving high performance of PSCs. In order to understand how GDYO affects the microscopic features of SnO₂ films, the dispersion solutions of the as-prepared SnO₂(GDYO) composite and SnO₂ were spin-coated on ITO, and the atomic

force microscopy (AFM) results are shown in Fig. 2a–d and Fig. S4 (ESI[†]). As shown in Fig. 2c, the microscopic features of the SnO₂(GDYO) film were significantly improved in comparison with those of the pure SnO₂ film (Fig. 2a). The root-mean-square roughness (RMS) decreased from 2.51 nm to 2.01 nm after 2% GDYO doping, implying the formation of a smoother film. Therefore, GDYO modification can directly improve the microscopic features of the SnO₂ surface, which will ultimately enhance the performance of PSCs.⁴² Fig. S5 (ESI[†]) shows the drop contact angles of the SnO₂ and SnO₂(GDYO) films. The slight change of the contact angle suggests that there is no significant difference in surface wettability. Furthermore, we performed top-view scanning electron microscopy (SEM) to examine the perovskite films on top of SnO₂ and SnO₂(GDYO) ETLs. As shown in Fig. 2e–h, it was noted that uniform, fully covered and hole-free perovskite layers with similar

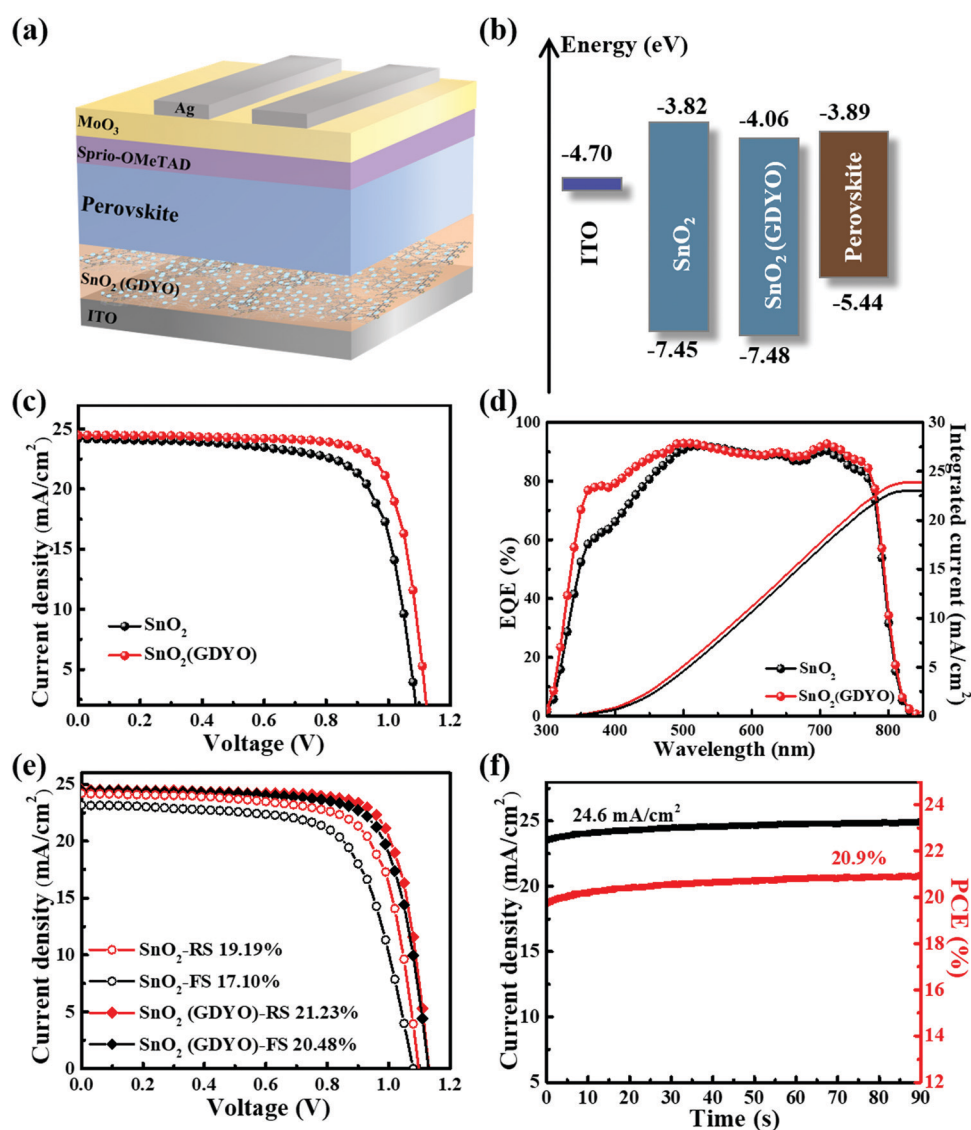


Fig. 3 (a) Schematic of the device structure of SnO₂(GDYO)-containing PSCs. (b) Diagram of the energy levels of the SnO₂, SnO₂(GDYO), and perovskite layers. (c) J–V curves, (d) EQE spectra and (e) J–V curves under reverse and forward scans of devices with and without GDYO modification. (f) Stabilized power output of the SnO₂(GDYO)-based device.

morphologies were obtained, but their grain sizes slightly differed. When the doping concentration of GDYO was 2%, the corresponding grain size reached its maximum. Consequently, the increase of grain size is beneficial to reducing grain boundary defects, thus suppressing non-radiative recombination and promoting carrier transport.⁴³

The structure of the SnO₂(GDYO)-based device, comprising ITO/SnO₂(GDYO)/perovskite/spiro-OMeTAD/MoO₃/Ag, is shown in Fig. 3a. The performance of perovskite devices is highly dependent on the optical properties of the adjacent ETLs. The transmission spectra of different ETLs were studied by UV-Vis spectroscopy. As shown in Fig. S6a (ESI[†]), the SnO₂ and SnO₂(GDYO) films show good light transmittance in the visible light range, which is advantageous for the perovskite film to absorb the incident light. Meanwhile, the band gaps (E_g) of these two films were calculated as 3.63 eV and 3.42 eV, respectively (Fig. S6b, ESI[†]). Normally, the energy level of ETL needs to be well matched with that of the perovskite layer, which determines the device performance, especially for V_{OC} . To examine this, we studied the energy bands of the GDYO-free and GDYO-doped ETLs using ultraviolet photoelectron spectroscopy (UPS). The cutoff energy regions and the valence band maximum energy regions of the SnO₂ and SnO₂(GDYO) composite layers are depicted in Fig. S6c and d (ESI[†]). As shown in Fig. 3b, the valence band maxima (E_V) of these ETLs are -7.45 eV (SnO₂) and -7.48 eV (SnO₂(GDYO)), respectively. Then, the corresponding conduction band minimum (E_C) can be determined by the equation of $E_C = E_V + E_g$. E_C values of -3.82 (SnO₂) and -4.06 eV (SnO₂(GDYO)) were obtained. Compared with the SnO₂ film, E_C of SnO₂(GDYO) increased by 0.24 eV, suggesting that SnO₂(GDYO) matches well with the energy level of the active layer; this will not only result in enhancement of the V_{OC} , but also promote the extraction and injection of electrons and consequently improve the short-circuit current density (J_{SC}) of the PSCs.⁴⁴

To evaluate the effect of SnO₂ ETLs with and without GDYO on the photovoltaic performance, ITO/SnO₂ or SnO₂(GDYO) ETL/Perovskite/Spiro-OMeTAD/MoO₃/Ag heterojunction cells were fabricated. Fig. 3c describes the corresponding J - V characteristics obtained under AM 1.5G irradiation (100 mW cm⁻²). The device with the pure SnO₂ ETL exhibited a PCE of 19.19%, with a J_{SC} of 24.20 mA cm⁻², a V_{OC} of 1.10 V, and a fill factor (FF) of 72.42%. In order to optimize device efficiency, varying GDYO doping concentrations were applied to SnO₂. As shown in Fig. S7 (ESI[†]), the best PCE was obtained when the doping concentration of GDYO was 2%. The resultant device treated with GDYO yielded a peak PCE of 21.23% with a V_{OC} of 1.13 V, a J_{SC} of 24.49 mA cm⁻², and an FF of 76.85%. Interestingly, it was found that the performance of all the GDYO-containing devices was superior to that of the control device. The better energy level arrangement and the reduction of defects in the SnO₂/perovskite interface mainly account for the improved device performance. The corresponding EQE spectra are plotted in Fig. 3d. It can be observed that the whole spectrum is elevated in the visible light range, indicating higher extraction efficiency of the SnO₂(GDYO)-based device (integrated current

23.91 mA cm⁻²) compared to that of the GDYO-free device (integrated current 23.06 mA cm⁻²). The EQE results of the SnO₂(GDYO)-based device and control device concur with the J - V curves. In addition, as shown in Fig. 3e, the hysteresis of the SnO₂(GDYO)-based device was obviously mitigated. The hysteresis index (HI) is used to evaluate the degree of hysteresis in accordance with the equation of $HI = (PCE_{Reverse} - PCE_{Forward})/PCE_{Reverse}$,⁴⁵ where $PCE_{Forward}$ and $PCE_{Reverse}$ represent the values of PCE during forward and reverse scanning, respectively. The HI values of devices based on the SnO₂ and SnO₂(GDYO) ETLs are 0.1 and 0.03, respectively. The hysteresis curve of J - V can be attributed to the presence of traps on the surface and the charge accumulation between the SnO₂ layer and perovskite layer, as widely reported in the literature. After adding GDYO, the hysteresis can be significantly suppressed by reducing the trap state and effective electron extraction in the perovskite layer.

Fig. 3f shows the best steady-state power and photocurrent output device measured at its maximum power point ($V_{MPP} = 0.85$ V). With the photocurrent stabilized at 24.6 mA cm⁻², the PCE of the SnO₂(GDYO)-based device can be retained at 20.9% for more than 90 s, consistent with the value extracted from the J - V curves. Besides, the GDYO-based device exhibited a narrow PCE distribution with a high average PCE, and the distribution of the other parameters fit a similar pattern (Fig. S8, ESI[†]).

We measured the charge transfer properties in order to understand the effect of GDYO on improving the device performance. The mechanism of device recombination can be understood by determining the relationship between the photoelectric response and the intensity of incident light. It can be reflected by plotting V_{OC} and J_{SC} as functions of incident light intensity according to the following equations:

$$J \propto \alpha I \quad (1)$$

and

$$V_{OC} = \frac{nk_B T \ln(I)}{q} + C \quad (2)$$

Here, I represents the light intensity, α is an ideal exponential factor related to charge extraction, k_B is the Boltzmann constant, T stands for the absolute temperature, q is the elementary charge, C is a constant, and n denotes an ideal factor related to recombination. The α values for the SnO₂(GDYO)-based device and control device were found to be 0.99 and 0.98, respectively (Fig. 4a). The slope of the GDYO-based device was slightly larger than that of the control device, indicating that the bimolecular recombination was inhibited at the SnO₂/perovskite interface because of the enhanced charge extraction and transfer ability. Fig. 4b shows the dependence of V_{OC} on $\log(I)$. Previous studies have shown that the value of n reflects trap-dominated monomolecular recombination. In this case, the n values of the pristine SnO₂ and SnO₂(GDYO)-based devices were 1.63 and 1.55, respectively, which indicates that GDYO assists in reducing monomolecular recombination and thus results in fewer defect states for the SnO₂(GDYO)-based device.⁴⁶

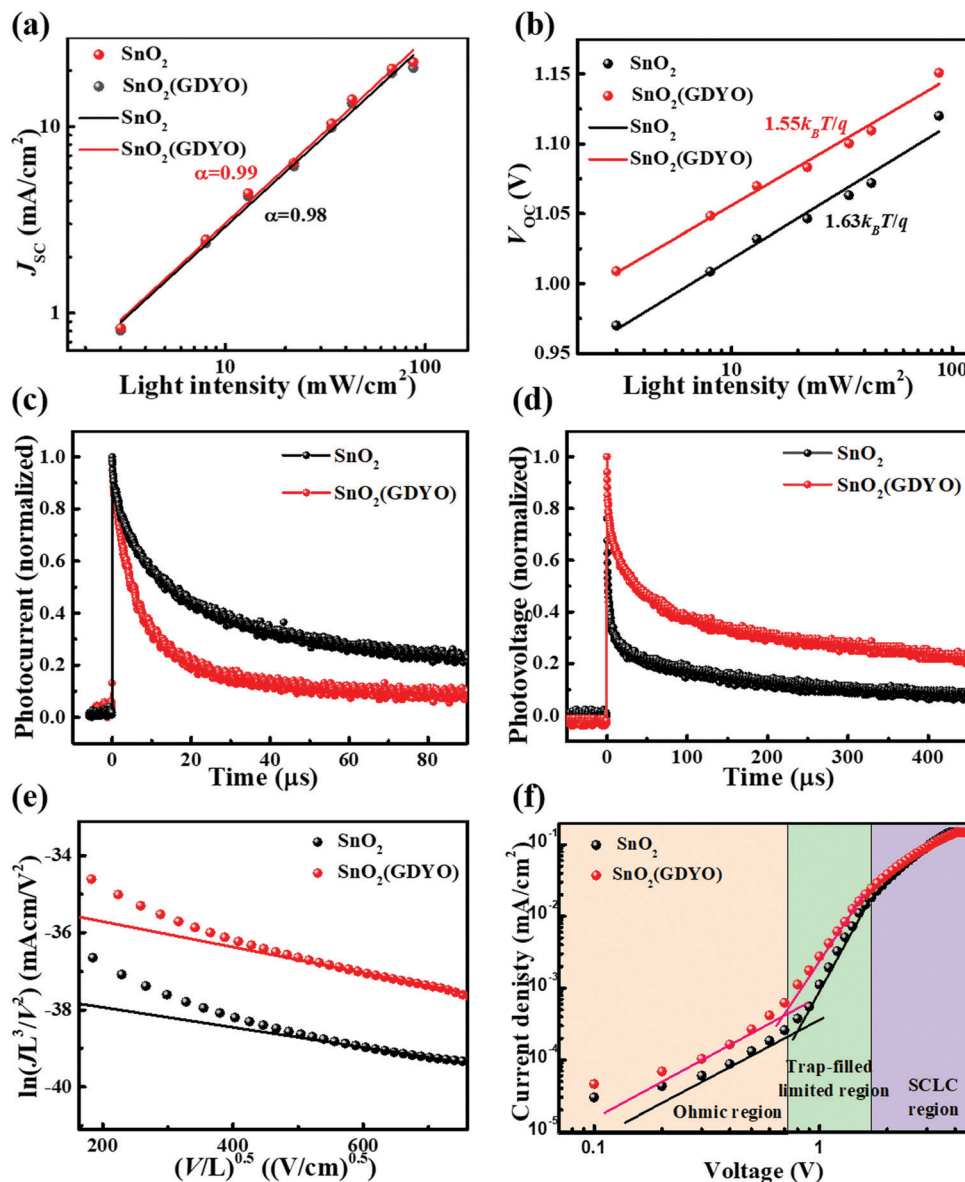


Fig. 4 (a) J_{SC} and (b) V_{OC} dependence upon different light intensities of PSCs based on the SnO_2 and $\text{SnO}_2(\text{GDYO})$ ETLs. (c) Transient photocurrent decay curves and (d) transient photovoltage decay curves of PSCs based on SnO_2 and $\text{SnO}_2(\text{GDYO})$ ETLs. (e) $J-V$ plotted in the format $\ln(JL^3/V^2)$ versus $(V/L)^{0.5}$ for the device structure of $\text{ITO}/\text{Ag}/\text{SnO}_2/\text{Ag}$ or $\text{ITO}/\text{Ag}/\text{SnO}_2(\text{GDYO})/\text{Ag}$. (f) Dark $J-V$ curves of electron-only devices based on the SnO_2 and $\text{SnO}_2(\text{GDYO})$ ETLs.

In order to investigate the charge extraction ability upon doping GDYO, we also measured the transient photocurrent decay (TPC) and photovoltage decay (TPV) of the devices. As shown in Fig. 4c, the photocurrent decay time (transport time constants τ_t) of the device for $\text{SnO}_2(\text{GDYO})$ was shorter than that of its counterpart ($15.9 \mu\text{s}$ vs. $29.3 \mu\text{s}$). Moreover, charge-recombination lifetime τ_r increased from $69.1 \mu\text{s}$ to $182.5 \mu\text{s}$ (Fig. 4d) after the addition of GDYO, which is indicative of more efficient charge extraction and reduced charge recombination.⁴⁷ To further explore the carrier motility of the SnO_2 and $\text{SnO}_2(\text{GDYO})$ films, we measured the space charge limited conduction (SCLC) by using electron-only devices with the structures of $\text{ITO}/\text{Ag}/\text{SnO}_2/\text{Ag}$ and $\text{ITO}/\text{Ag}/\text{SnO}_2(\text{GDYO})/\text{Ag}$ (Fig. 4e). The electron mobility of ETL films can be obtained by the field-independent Mott-Gurney equation:

$$J = \frac{9}{8} \epsilon_0 \epsilon_r \mu \frac{V^2}{d^3}$$

where J is the current density, ϵ_0 is the permittivity of the vacuum, ϵ_r is the relative permittivity of SnO_2 , μ is the zero-field mobility of electrons, d is the thickness of the ETL, and V represents the applied potential to the device. Accordingly, the electron mobility of the pure SnO_2 -based electron-only device was $8.18 \times 10^{-4} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, which is apparently lower than that of the $\text{SnO}_2(\text{GDYO})$ -based counterpart ($8.15 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$). We attribute the increase of electron mobility to the highly electronic conductivity of GDYO. High electron mobility can accelerate electron transport which is critical to increase current.⁴⁸

In addition, we conducted electrochemical impedance spectroscopy (EIS) analysis, as shown in Fig. S9 (ESI[†]). The diameter of a semicircle depicts the charge recombination

resistance (R_{rec}), which is related to the leakage current and charge recombination of the perovskite layer and ETL. The $\text{SnO}_2(\text{GDYO})$ -based device displayed a larger semicircle diameter than the control device, suggesting that less current leakage and electron recombination occurred at its surface. As a result, this facilitates the prevention of charge accumulation at the interface and improves the electron transfer efficiency.⁴⁹

Electron-only devices with the structure of ITO/ETLs/perovskite/PCBM/Ag were fabricated in order to study the trap-state density at the interface of the perovskite and ETLs. Fig. 4f shows the corresponding J - V curves under dark conditions. The fitting curve consists of three regions: the ohmic region at low bias voltage, the trap-filling limited region, and the trap-free SCLC region at high voltage. The trap-state density N_t can be calculated by the trap-filled limit voltage using the following equation:

$$N_t = \frac{2\epsilon\epsilon_0 V_{\text{TFL}}}{el^2} \quad (3)$$

where e is the electron charge, l is the thickness of perovskite, V_{TFL} is the trap-filled limit voltage, ϵ_0 represents the vacuum permittivity, and ϵ is the relative dielectric constant of perovskite. Compared to the control device without GDYO, the device with the $\text{SnO}_2(\text{GDYO})$ ETL showed a smaller V_{TFL} . The trap-state density for the $\text{SnO}_2(\text{GDYO})$ and SnO_2 -based devices were determined as $6.48 \times 10^{16} \text{ cm}^{-3}$ and $7.85 \times 10^{16} \text{ cm}^{-3}$, respectively. Apparently, the defect density decreased after using GDYO to

modify the SnO_2 film, which suggests that GDYO can drastically passivate the trap-state density in the perovskite film as well as the interface between the ETL and perovskite.⁵⁰ Besides, the aforementioned SEM results prove the increase of the grain size of the perovskite layer after doping with GDYO, which contributes to the decrease of the defect density.

Steady state photoluminescence spectroscopy (PL) was tested to further analyze the efficacy of doping GDYO in performance enhancement. Fig. 5a shows the photoluminescence spectra of a pristine perovskite film and the perovskite film deposited on SnO_2 film and $\text{SnO}_2(\text{GDYO})$ film, respectively. We measured the PL spectrum with an excitation light of 630 nm. The photoluminescence emission from perovskite at 790 nm is consistent with previous reports. Generally, a strong PL quenching indicates a high separation rate of photoexcited electron-hole pairs. Comparing the intensity of all peaks, a pronounced quenching effect was observed when GDYO was added to SnO_2 . This indicates that the $\text{SnO}_2(\text{GDYO})$ -based device demonstrated higher electron extraction capability, which facilitates electron transport from perovskite to the ETL. The fast electronic transport capability can reduce the accumulation and charge recombination at the interface, which explains the enhancements in J_{sc} and FF.⁵¹

The time resolved photoluminescence (TRPL) was measured with the same device structure in order to elucidate the dynamic process of electron transfer and extraction. As shown

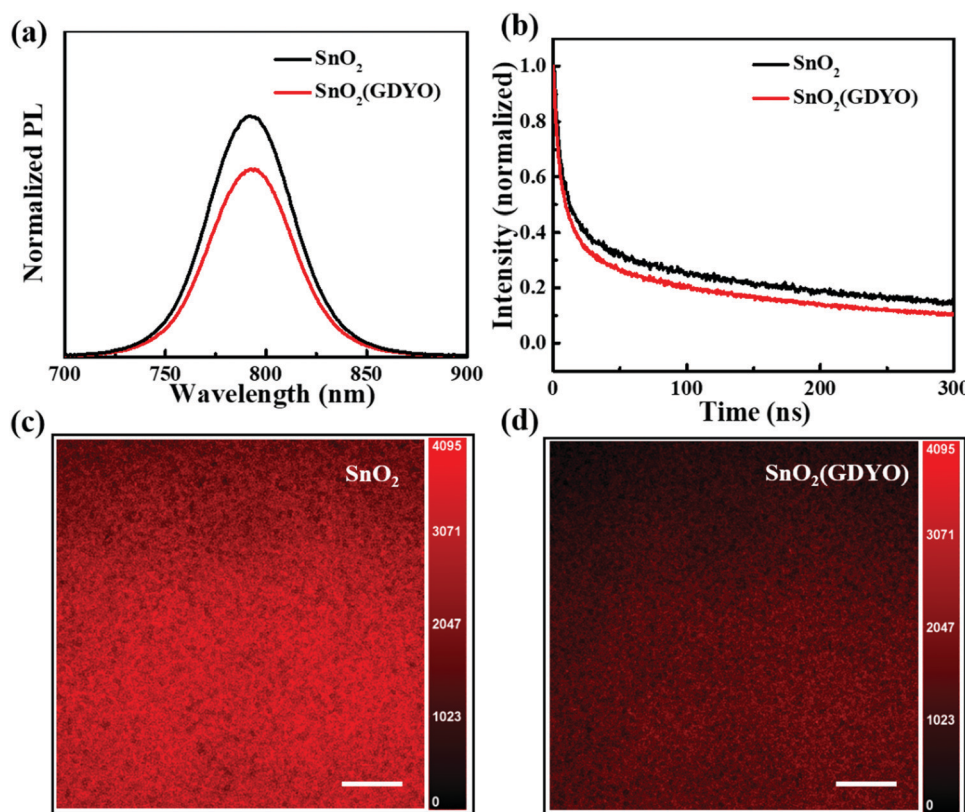
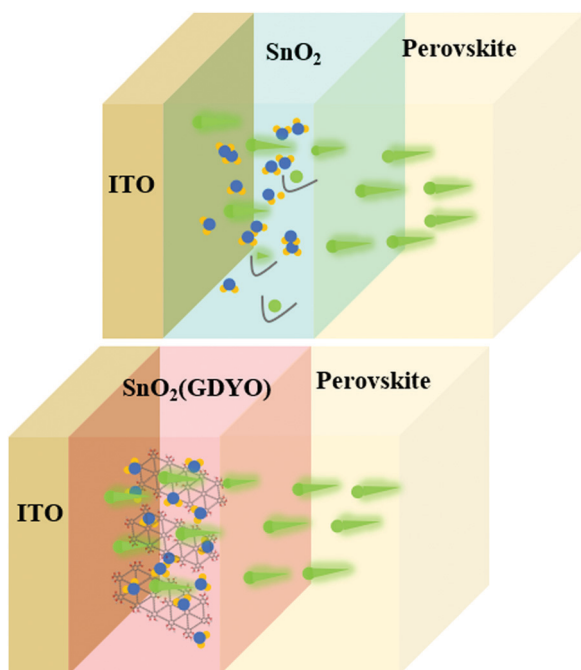


Fig. 5 (a) Steady-state and (b) time-resolved photoluminescence spectra of perovskite films deposited on different substrates. Confocal fluorescence microscopy of perovskite films deposited on the (c) SnO_2 ETL and (d) $\text{SnO}_2(\text{GDYO})$ ETL; the scale bar is 1 μm .

Table 1 The values of the decay amplitude constants and decay times

Samples	τ_1 (ns)	A_1	τ_2 (ns)	A_2	τ_{ave} (ns)
SnO ₂	12.59	0.394	344.9	0.482	335.3
SnO ₂ (GDYO)	10.44	0.619	287.4	0.297	266.6

in Fig. 5b, the data were fitted with two exponential decay curves to yield the lifetime of the carriers according to the equation of $f(t) = A_1 \exp(-t/\tau_1) + A_2 \exp(-t/\tau_2)$. The fast decay process (τ_1) is mainly composed of bimolecular recombination of photo-generated free carriers, and the slow decay process (τ_2) is produced by trap-assisted recombination; A_1 and A_2 represent the decay amplitude. Ultimately, the average decay time can be calculated, and the detailed parameters are listed in Table 1. The lifetimes obtained from TRPL measurements were 335.3 ns and 266.6 ns for the SnO₂ and SnO₂(GDYO) films, respectively. It was observed that the SnO₂(GDYO) film showed a reduced PL lifetime, indicating that GDYO effectively enhanced the electronic extraction and transport capabilities of SnO₂.⁵² Aiming to intuitively understand the carrier quenching dynamics, confocal fluorescence microscopy was used for characterization. As can be seen from Fig. 5c and d, the perovskite deposited on the SnO₂(GDYO) film appeared darker than that on the SnO₂ film, which is attributed to the efficient transport of carriers.⁵³ This indicates that more carriers are transported to the transport layer instead of quenching on the perovskite surface after GDYO modification, which is consistent with the PL results. Fig. 6 shows the effect of GDYO doping on the electron transfer. There are many surface defects in the SnO₂ ETL, which are mainly caused by oxygen vacancy defects formed by uncoordinated Sn on the surface. These defects can trap electrons and become the recombination centers of photogenerated carriers,

**Fig. 6** Schematic of the effect of GDYO doping of SnO₂.

thus seriously affecting the charge extraction and transport. GDYO can passivate these defects and promote carrier transport, which plays a positive role in improving the J_{SC} and FF of the device and thus the overall performance of the device.

The thermal and illumination stability of the SnO₂ and SnO₂(GDYO)-based devices were tested with poly(triarylamine) (PTAA) replacing spiro-OMeTAD as the hole transport material. The initial PCEs of the SnO₂ and SnO₂(GDYO)-based devices with a PTAA hole transport layer were 15.1% and 17.6%, respectively (Fig. S10, ESI[†]). As shown in Fig. S11a (ESI[†]), the thermal stability test was carried out in a nitrogen box with heating at 80 °C; over 80% of the initial value of the GDYO-based device was retained after 25 days, which was much better than the control device. We also tracked the light stability under 1 sun illumination (25 °C, under nitrogen storage), and the GDYO-based device also showed better stability (Fig. S11b, ESI[†]). The stability tests indicate that GDYO modification of a SnO₂ ETL is an effective approach to improve device stability. The interaction between GDYO and SnO₂, associated with reduced oxygen vacancy defects, improved the carrier transport efficiency and reduced the charge recombination, consequently contributing to the improved device stability.

3. Conclusion

To conclude, we have introduced an effective interface engineering approach for a high-performance planar PSC by using a novel SnO₂(GDYO) ETL. It is demonstrated that the terminal carboxy and hydroxy groups of GDYO can link the uncoordinated Sn in SnO₂ *via* chemical bonds, leading to reduced oxygen vacancy of SnO₂; this further facilitates the suppression of the non-radiative recombination and promotes the charge transport. Upon the incorporation of GDYO, the quality of both the SnO₂ film and perovskite film was improved. As a result, the SnO₂(GDYO)-based device achieved a high PCE of 21.23%, which is an enhancement of over 10% as compared to the control device. Furthermore, the device stability was improved. The SnO₂(GDYO)-based device maintained 71% of its initial PCE after 160 h under continuous illumination and 84% after 24 days of heating at 80 °C. These findings highlight the significance of precise interface engineering toward high-performance PSCs and extend the promising applications of functionalized GDY in photovoltaic technologies.

Conflicts of interest

There are no conflicts of interest to declare.

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