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A chelate like no other: exploring the synthesis, coordination chemistry and applications of imidoyl amidine frameworks

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Since the remarkable work by Werner, the design and use of chelating ligands have played an essential part in coordination chemistry. While ligands such as 2,4-pentanedione (acac) and 1,5-diazapentadienyl (NacNac) have been established in the literature for over a century and their rich coordination chemistry is well known, 1,3,5-triazapentadiene ligands remain the significantly less explored members of this family. Also known as imidoyl amidine (**ImAm**), the coordination chemistry of this ligand framework has recently attracted a great deal of attention due to its multiple binding sites, ability to be incorporated into organic and inorganic materials, and diverse applicability (catalysis, biomolecular probes, materials science, etc.) making them a class of chelates like no other. The scope of this review is to provide a critical overview of the research progress on the chemistry of **ImAm** as well as complexation with different metal ions. General synthetic routes and new insights for the preparation of **ImAm** are discussed along with their use as starting materials or intermediates in organic and coordination chemistry in the development of materials that have found use in fields ranging from magnetic, conducting and luminescent materials to catalysis and biomedical applications. Concluding remarks including the limitations and perspectives of **ImAm** are discussed.

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1. Introduction

Classical β -diketones or 1,3-diketones ($R_1COC(R_2)COR_1$) represent one of the most valuable classes of ligands for metal ion chelation (Scheme 1). They have been studied for more than a century and their rich coordination chemistry is well explored

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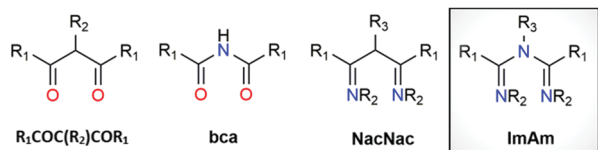
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Niki Mavragani

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Scheme 1 Structure of 1,3-diketones ($R_1COC(R_2)COR_1$), bis(carbonyl)amide (bca), 1,3-diketoiminato (NacNac) and imidoil amidine (ImAm).

along with their industrial applications.¹ Substitution of the central α -carbon in $R_1COC(R_2)COR_1$ for nitrogen affords the bis(carbonyl)amide ligand (bca; Scheme 1), which has been employed in coordination complexes for a variety of applications in materials chemistry including single-chain magnets, spin crossover (SCO) materials, metal–organic frameworks (MOFs) and luminescent complexes.² Since the earliest development of the coordination chemistry of dinitrogen analogues of 1,3-dicarbonyl compounds (*i.e.*, 1,5-diazapentadienyl, β -diketiminates or NacNac) by Bradley³ and Holm,⁴ a great deal of research has focused on the use of NacNac leading to it becoming one of the most employed monoanionic nitrogen-based ligands in the field (Scheme 1).⁵ Indeed, NacNac ligand frameworks have been much less explored compared to the $R_1COC(R_2)COR_1$ analogues but their complexes show various applications in many fields.⁶

The less explored member of this family (compared to $R_1COC(R_2)COR_1$, bca and NacNac) is the trinitrogen analogue of the 1,3-dicarbonyls, *i.e.*, 1,3,5-triazapentadiene or imidoil amidine (Scheme 1). While both terms are equally applied in literature, for the purposes of this article the latter (imidoil amidine) will be used exclusively with the corresponding abbreviation **ImAm**. The presence of an additional N-donor atom (compared to NacNac) in **ImAm** has triggered the attention of coordination chemists. In addition to the *N-N* bidentate chelate coordination mode (also known as U-shape fashion)

of NacNac, **ImAm** ligands can also coordinate in a tridentate fashion (W-shape) when donor atoms are incorporated onto the substituents at the β -carbon atoms or act as linkers between metal ions in their deprotonated/anionic form.⁷ The ability of these frameworks to adopt multiple coordination modes (*vide infra*) forming stable five- and six-membered chelate rings, can be exploited to efficiently target coordination to various metal ions including heterometallic systems. Moreover, the central N-atom can easily switch between its protonated and deprotonated form based on the pH conditions and act as an “on-off” switch, particularly attractive for luminescent applications.⁸ Computational studies of trinuclear heterometallic coordination complexes of Fe^{II} with Zn^{II} , Ni^{II} , Cu^{II} or Co^{II} , based on **ImAm** ligands with pyrazole functional groups, revealed that weak exchange interactions are expected to occur between the metal centers, undergoing a two-step SCO process.⁹ Another interesting feature of the **ImAm** framework is the ability to functionalize the carbon atoms (R_1) and/or the nitrogen atoms (R_2 and R_3), leading to a vast array of organic molecules with desirable properties. Despite these advantages and the potential of **ImAm** ligands, complexes containing a pre-synthesized **ImAm** are rather rare compared to other 1,3-dicarbonyl derivatives and analogues. This can be attributed to the challenge associated with isolating non-coordinated **ImAm** frameworks, which until recently²¹ the synthesis was poorly developed.

Having established facile synthetic methodology for the isolation of **ImAm** chelates and their derivatives, enables one to tap into the potential of this versatile framework. To that end, the scope of the present review is to provide new insights into the preparation of **ImAm** along with its usage as building blocks in organic chemistry and employment as ligands in coordination chemistry. Our aim is to cover the recent research progress on the chemistry and applications of **ImAm**, excluding biguanidine and related derivatives as they have already been reviewed¹⁰ and will therefore not be addressed herein.



Muralee Murugesu

the University of Ottawa as an assistant professor and since 2015 he is a full professor. Prof. Murugesu's research focuses on the design and development of single-molecule magnets, metal–organic frameworks and high-energy materials.

Muralee Murugesu received his PhD from the University of Karlsruhe in 2002 under the supervision of Prof. A. K. Powell. He undertook postdoctoral stays at the University of Florida (2003–2005) with Prof. G. Christou, and jointly at the University of California, Berkeley and the University of California, San Francisco under the supervision of Prof. J. R. Long and the Nobel Laureate Prof. S. Pruisner (2005–2006). In 2006, he joined



Jaclyn L. Brusso

Jaclyn Brusso obtained her PhD from the University of Waterloo under the direction of R. T. Oakley in 2006. Following an NSERC Postdoctoral Fellowship with D. Perepichka at McGill University and then with S. Holdcroft's group at Simon Fraser University, Jaclyn joined the University of Ottawa in 2010. Jaclyn's research program focuses on the design and development of novel organic/inorganic based optical, magnetic and electronic materials.



2. Brief introduction to the organic chemistry of the imidoyl amidine frameworks

2.1 Tautomerism of imidoyl amidines

Imidoyl amidines are multifunctional, nitrogen-rich molecules bearing formally fused amidine, imide and amine groups forming a characteristic unsaturated N–C–N–C–N chain. As such, **ImAm** can exist in two main tautomeric forms, namely 1,3,5-triaza-1,3-pentadienes (*i.e.*, amino–imino; Scheme 2a) and 1,3,5-triaza-1,4-pentadienes (*i.e.*, imino–imino; Scheme 2b), depending on the position of the double bond. Upon coordination of deprotonated **ImAm** ligands to a metal center, delocalization of the double bond system over the entire molecule occurs¹¹ while the C–N bonds are usually localized in the neutral form.¹² Based on the number of hydrogen atoms attached to the nitrogen atoms, **ImAm** can be classified as primary (three N–H groups), secondary (two N–H groups), tertiary (one N–H group) or quaternary (where no N–H groups are present).

2.2 Synthetic approaches for the development of imidoyl amidines

The majority of literature references that contain **ImAm** involves their preparation through metal assisted reactions (*vide infra*) with only a few examples of the synthesis of such organic molecules without metal mediated transformations.^{13–21} Thus far, reported methods for the preparation of **ImAm** chelates mostly lead to heavily *N*-substituted **ImAm** with bulky substituents at the terminal nitrogen atoms and/or strong electron-acceptor functional groups at the carbon atoms, as shown in Scheme 3.¹³ Since these synthetic routes have already been discussed in detail by Kopylovich and Pombeiro,¹⁴ herein we will only focus on the main advantages and disadvantages of each of these procedures. For example, the preparation of non-coordinated **ImAm** via the Pinner reaction¹⁵ (Scheme 3a) leads to rather low yields and a large number of by-products, making this a relatively inefficient preparative route. A more versatile and simplified method proposed by Ley and Muller¹⁶ (Scheme 3b) leads to relatively good yields and provides a pathway towards the isolation of a large number of **ImAm** (where R₁, R₂, R₃ can be either phenyl groups or different substituted phenyl groups). As highlighted in Scheme 3c step I, *N*-imidoylimidoates can be obtained by the reaction of imidoate hydrochlorides with *N*-imidoyl chlorides under basic conditions and, upon further reaction with a primary or secondary amine (Scheme 3c



Scheme 2 The 1,3,5-triaza-1,3-pentadienes or amino–imino (a) and 1,3,5-triaza-1,4-pentadienes or imino–imino (b) tautomeric forms of **ImAm** ligands.

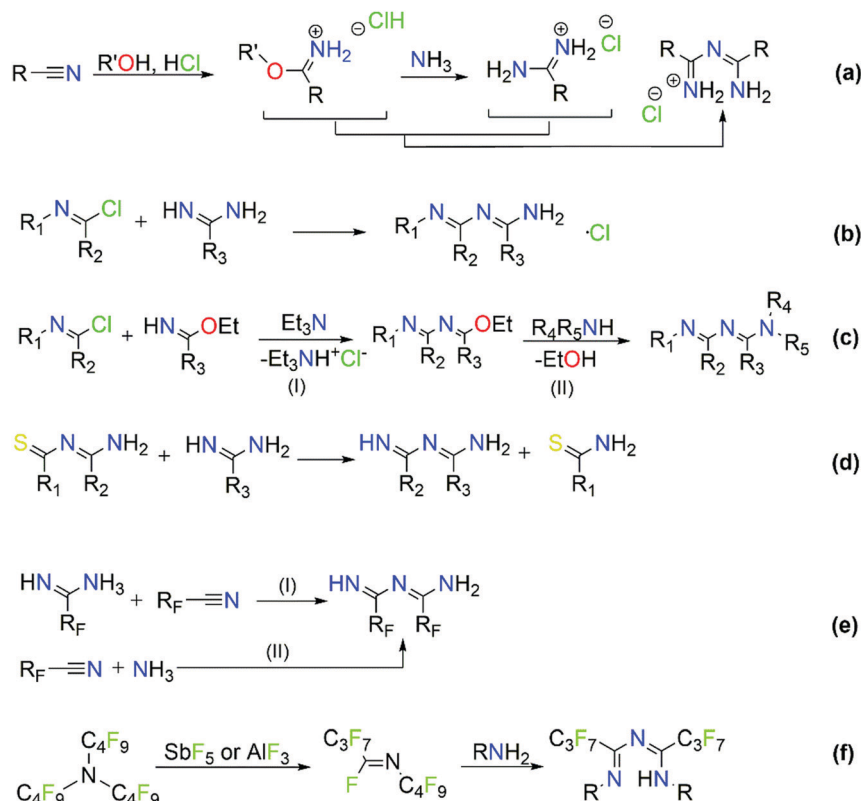
step II), tertiary or quaternary **ImAm** frameworks can also be isolated. This was well demonstrated by Häger *et al.* for the synthesis of tertiary and quaternary **ImAm** ligands and their complexation with various transition metal ions as well as boron difluoride.^{13c} Alternative synthetic routes involve the preparation of *N*-tobenzoylbenzamidines proposed by Titherley and Hughes,¹⁷ which are then converted to the corresponding **ImAm** by the desulfurization reaction with an amine or amidine (Scheme 3d).¹⁵ Condensation of perfluorocarbon (or perchloro-) activated amidines with similarly activated nitriles (Scheme 3e, reaction I) or by the reaction of a perfluoroalkynitrile with an excess of anhydrous ammonia (Scheme 3e, reaction II) affords halide substituted **ImAm**.¹⁸ The main disadvantage of these substituted **ImAm** frameworks is their rather unstable nature, which leads to cyclization thereby converting them to the corresponding triazines.¹⁹ An alternative procedure for the synthesis of **ImAm** containing perfluoro moieties involves the reaction of perfluoro-5-aza-4-nonene with primary amines (Scheme 3f).²⁰ This synthetic route gives access not only to halide substituted **ImAm** but also to *N*-substituted **ImAm**, depending on the initial amine used for the synthesis.

As mentioned, these preparative routes typically afford heavily *N*-substituted **ImAm** with bulky substituents at the terminal nitrogen atoms and/or strong electron-acceptor functionalities at the carbon atoms. Additionally, the overall yields reported are poor and, in most cases, the starting materials are not commercially available, thus adding more steps to their synthesis. Recently, we developed a high yielding one-pot metal free synthetic procedure for the preparation of carbon substituted **ImAm**.²¹ This procedure gave access to the isolation of *N*-2-pyridylimidoyl-2-pyridylamidine (**Py₂ImAm**) and *N*-2-pyrimidylimidoyl-2-pyrimidylamidine (**Pm₂ImAm**), starting from commercially available 2-pyridinecarbonitrile and 2-pyrimidinecarbonitrile, respectively (Scheme 4). The synthesis of both compounds involves the reaction of ammonia gas with the appropriately substituted nitriles. In the case of **Py₂ImAm**, the reaction is carried out under increased pressure affording a 63% yield. The **Pm₂ImAm** is isolated by allowing 2-pyrimidinecarbonitrile to react slowly with ammonia over a period of two weeks at room temperature and under atmospheric pressure resulting in a 76% yield. Encouragingly, the versatility of this procedure enables the preparation of various **ImAm** in high yields depending on the substituted nitriles in use. With that said, attempts to synthesize *N*-2-thienylimidoyl-2-thienylamidine (**Th₂ImAm**) by reacting 2-cyanothiophene with NH₃(g) led to negligible yields of the desired product regardless of the temperature or pressure conditions used.^{35d} This was overcome by treating 2-thienylamidine with sodium hydride thereby strengthening the nucleophilicity of the amidine, which subsequently reacts with 2-cyanothiophene to form **Th₂ImAm** in high yields and purity (Scheme 4).

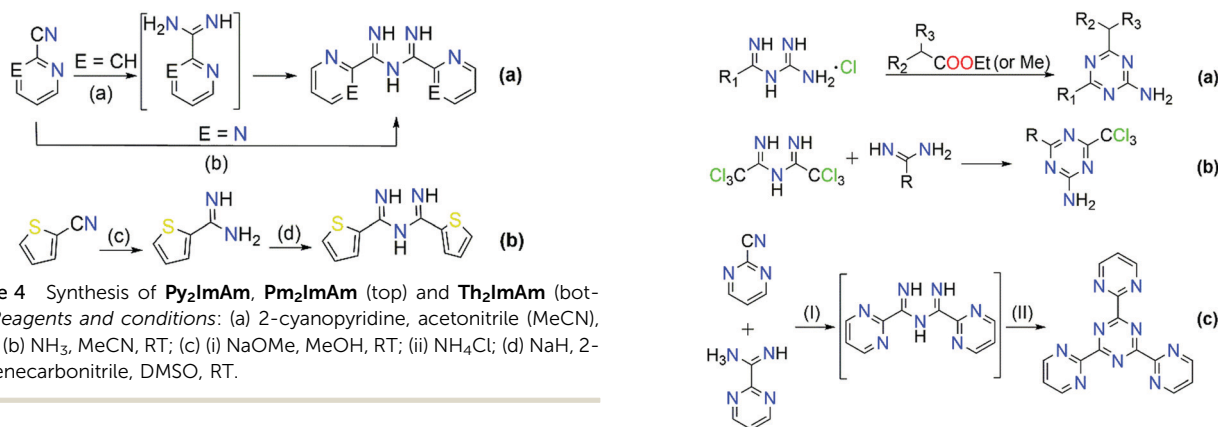
2.3 Imidoyl amidines as building blocks in organic materials

Although their synthesis has not been widely investigated, **ImAm** moieties have attracted interest as building blocks in the development of new organic frameworks. For example, **ImAm**





Scheme 3 Common synthetic routes for the synthesis of non-coordinated **ImAm** frameworks.



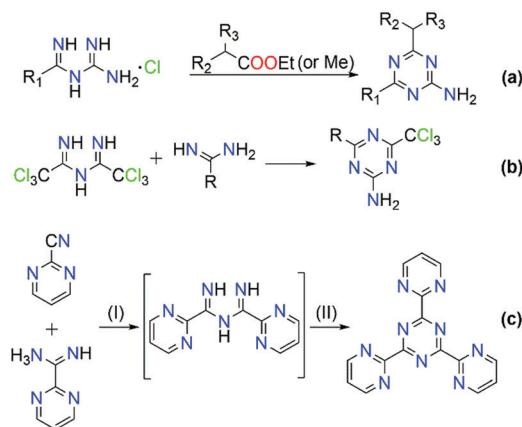
Scheme 4 Synthesis of **Py₂ImAm**, **Pm₂ImAm** (top) and **Th₂ImAm** (bottom). Reagents and conditions: (a) 2-cyanopyridine, acetonitrile (MeCN), 110 °C; (b) NH₃, MeCN, RT; (c) (i) NaOMe, MeOH, RT; (ii) NH₄Cl; (d) NaH, 2-thiophenecarbonitrile, DMSO, RT.

have been successfully used as intermediates in the formation of triazines¹⁹ and as starting materials for the synthesis of aminotriazines (Scheme 5).²² In the past few decades these two classes of compounds have been extensively studied in various fields, such as medicinal chemistry due to their excellent antitumor and anti-inflammatory activity,^{22,23} in coordination and materials chemistry²⁴ as ligands for the preparation of Metal Organic Frameworks (MOFs) with significant CO₂ capture capabilities,²⁵ as well as in catalysis.²⁶

Additionally, **ImAm** moieties have been successfully used as starting materials for the preparation of boratriazines (BTA). Complexation of **ImAm** with boron leads to a new class of materials, similar to those of boron dipyrromethene (BODIPY),

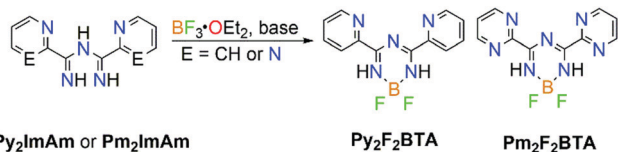
which have been studied for their sharp fluorescence peaks,²⁷ high thermal resistance,²⁸ strong UV-absorption²⁹ and excellent stability.²⁹ Recently, the incorporation of boron into **ImAm** frameworks was reported, yielding compounds which may be considered as unfused analogues to the extensively studied BODIPY dyes.^{21b} In Scheme 6, the general synthetic route for the synthesis of 2,2-difluoro-4,6-bis(2-pyridyl)-1,3-dihydro-1,3,5,2-triazaborinine (**Py₂F₂BTA**) and 2,2-difluoro-4,6-bis(2-pyrimidinyl)-1,3-dihydro-1,3,5,2-triazaborinine (**Pm₂F₂BTA**)

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Scheme 5 Synthesis of triazines and aminotriazines using **ImAm** as starting material. Reagents and conditions on (c): (i) NH₃, acetonitrile (MeCN), 110 °C, 72 h; (ii) MeCN, 110 °C, 120 h.

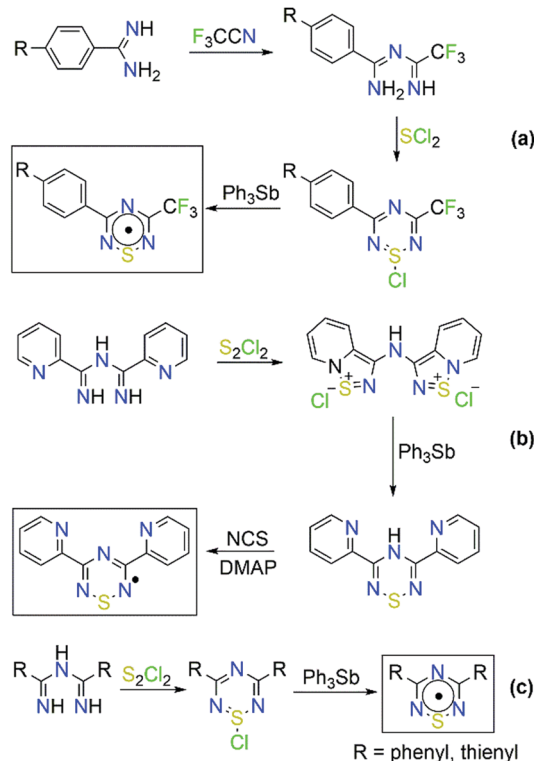




Scheme 6 General synthetic route used for the synthesis of 2,2-difluoro-4,6-bis(2-pyridyl)-1,3-dihydro-1,3,5,2-triazaborinine (**Py₂F₂BTA**) and 2,2-difluoro-4,6-bis(2-pyrimidinyl)-1,3-dihydro-1,3,5,2-triazaborinine (**Pm₂F₂BTA**). Reagents and conditions: (**Py₂F₂BTA**) triethylamine (Et₃N), tetrahydrofuran (THF), 60 °C; (**Pm₂F₂BTA**) Et₃N, dichloromethane (DCM).

is illustrated. Furthermore, the tridentate (*NNN*) coordination site in **Py₂F₂BTA** and **Pm₂F₂BTA** is expected to promote coordination through the chelating effect to a metal center.^{30a} Taking advantage of the ligating properties of these boratriazines, a series of **Py₂F₂BTA** and **Pm₂F₂BTA** coordination complexes with iron and cobalt salts that exhibited interesting magnetic properties (*i.e.* strong ferromagnetic interactions and field induced Single-Molecule Magnet (SMM) behavior) were reported.³⁰ As shown, it is possible to maintain the chelating ability of these organic ligands while tuning the physical properties of the construct. This approach may give access to interesting photocatalytic activity thanks to the photoactive features of the system and careful choice of the metal ion.

Organic radicals represent another field in which **ImAm** have been employed as starting materials, in particular for the synthesis of thiatriazinyls (TTA).³¹ In 1984, Oakley and co-workers first isolated and structurally characterized 3,5-bis(phenyl)-1,2,4,6-thiatriazinyl (**Ph₂TTA**) triggering the scientific interest for the preparation of such molecules.^{31b} Commonly, 1,2,4,6-thiatriazinyls can be prepared from the reaction of trichlorocyclotrihiazene with amidines³² or from the reaction of **ImAm** with excess SCl₂.³³ Boéré and coworkers proposed a general synthetic route for the synthesis of asymmetrical 5-aryl-3-trifluoromethyl-1,2,4,6-thiatriazinyls where **ImAm** treated with SCl₂ affords *S*-chlorothiatriazines, which can then be reduced with Ph₃Sb yielding the corresponding radicals (Scheme 7).³⁴ Recently, we developed the synthetic methodology for the preparation of symmetrical TTA such as 3,5-dipyridylthiatriazinyl (**Py₂TTA**) and bis-(2-thienyl)-1,2,4,6-thiatriazinyl (**Th₂TTA**).³⁵ In the case of **Py₂TTA**, reaction of **Py₂ImAm** with S₂Cl₂ affords the thiadiazolium dication, which upon treatment with Ph₃Sb undergoes a two-electron reduction affording **Py₂TTAH**. Upon oxidation of **Py₂TTAH** with *N*-chlorosuccinimide (NCS), in the presence of base, the radical form of **Py₂TTA** is isolated (Scheme 7). In their crystal packing, **Py₂TTA** and **Th₂TTA** form dimeric species *via* S··S contacts at ~2.6 Å. Electron paramagnetic resonance (EPR) spectroscopic studies on symmetrical and asymmetrical TTA revealed that the unpaired spin density is localized on the heterocyclic TTA ring. Additionally, cyclic voltammetric studies suggested that there is a monomer–dimer equilibria in solution.^{34,35} Lastly, the coordination chemistry of **Py₂TTA** has been explored, affording complexes with transition metal ions and lanthanides that exhibit interesting magnetic properties (*i.e.* SMM behavior), with the ligand in either the anionic or oxidized form.³⁶



Scheme 7 General synthetic routes for the synthesis of (a) asymmetrical 5-aryl-3-trifluoromethyl-1,2,4,6-thiatriazinyls, (b) symmetrical 3,5-dipyridylthiatriazinyl (**Py₂TTAH**) and 3,5-bis(phenyl)-1,2,4,6-thiatriazinyl (**Ph₂TTA**), and (c) 3,5-bis-(2-thienyl)-1,2,4,6-thiatriazinyl (**Th₂TTA**).

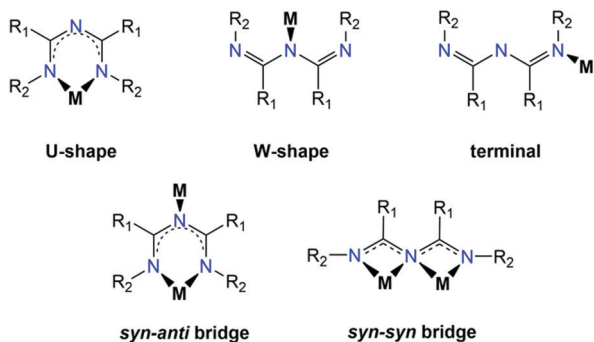
2.4 Metal-assisted formation of imidoyl amidine ligands

From a historic perspective, the first **ImAm** complex was reported in 1984.³⁷ The methylimidoyl-methylamidine ligand was generated *in situ* *via* a metal assisted transformation, when acetamidine was reacted in the presence of nickel metal ions. Subsequent work led to other metal assisted transformations to afford **ImAm** frameworks, such as the metal-mediated hydrolytic decomposition of oximes^{38a,b} or triazines,^{38c,d} the nucleophilic addition to dicyanamide,^{38e} the template condensation of nitriles and/or amidines at a metal center,^{38f-i} as well as *in situ* reactions of a metal ion and a nitrile under solvothermal conditions.^{38j} In 2011, a review by Kopylovich and Pombeiro¹⁴ covered this topic in an excellent way. Thus, all these synthetic approaches will not be described herein in detail. In the following sections, coordination complexes and significant progress in the field of the last decade will be covered.

3. The coordination chemistry of symmetrical imidoyl amidine chelates

While the coordination chemistry of **ImAm** is significantly influenced by the R₁, R₂ and R₃ substituents, functionalization at the central nitrogen atom (*i.e.*, R₃) is rare with this site either protonated or involved in coordination to transition metal ions. The R₁ and R₂ groups, on the other hand, can either contribute





Scheme 8 The crystallographically established coordination modes of **ImAm**.

to the coordination sphere of the metal ion by having donor atoms such as oxygen, nitrogen or sulfur, or not contribute at all. Furthermore, the presence of aryl-, aromatic or halogen-rich moieties are likely to stabilize discrete coordination complexes by forming stronger supramolecular networks (*via* H-bond, π - π interactions, halogen bonds, *etc.*). In general, there are three coordination modes that most commonly appear in the literature (Scheme 8). The U-shaped coordination mode allows for the **ImAm** ligand to act as a bidentate chelate forming a six-membered ring, which represents the most common coordination mode of **ImAm** based complexes and has yielded a large number of mononuclear complexes. On the other hand, **ImAm** can also act as ligands either in the W-shape (through the middle N atom) or terminal mode (through the peripheral N atom), leading mostly to mononuclear complexes. By utilizing both the U- and W-shape coordination modes, the *syn-anti* and *syn-syn* bridging motifs have led to the isolation of polynuclear complexes. The main difference between these two coordination modes is that the middle nitrogen atom in the case of the *syn-syn* bridging mode coordinates to two metal centers, while in the case of *syn-anti*, it binds to only one metal ion.

3.1 **ImAm** chelates without additional donor atoms in the substituents

In Scheme 9, **ImAm** chelates with no additional coordinating groups such as H, Me, Ph, CF_3 , C_6F_5 , C_3F_7 , *etc.* are presented. The synthesis of the dimethyl **ImAm** ligand (L^1) was achieved *in situ* in the presence of a Ni^{II} metal center.³⁹ Although deprotonation of the central N-atom can be relatively facile, depending on the reaction conditions, it is not always achieved. This was aptly illustrated in the work of Thiele *et al.* where they obtained the mononuclear cationic complex $[\text{Ni}^{\text{II}}(\text{L}^1)_2]\text{Cl}$ (**1**) in which one of the L^1 ligands was neutral while the other was anionic (Fig. 1).^{39b} By changing the metal source from NiCl_2 to NiSO_4 enabled isolation of a mononuclear Ni^{II} complex, where both ligands were in their neutral form.^{39b} In both complexes, the ligand acts as a chelate stabilizing the 3d metal ion in a slightly distorted square planar geometry. The same group also reported a similar square planar mononuclear Ni^{II} complex with *in situ* preparation of L^2 in which both ligands are deprotonated.^{39b}

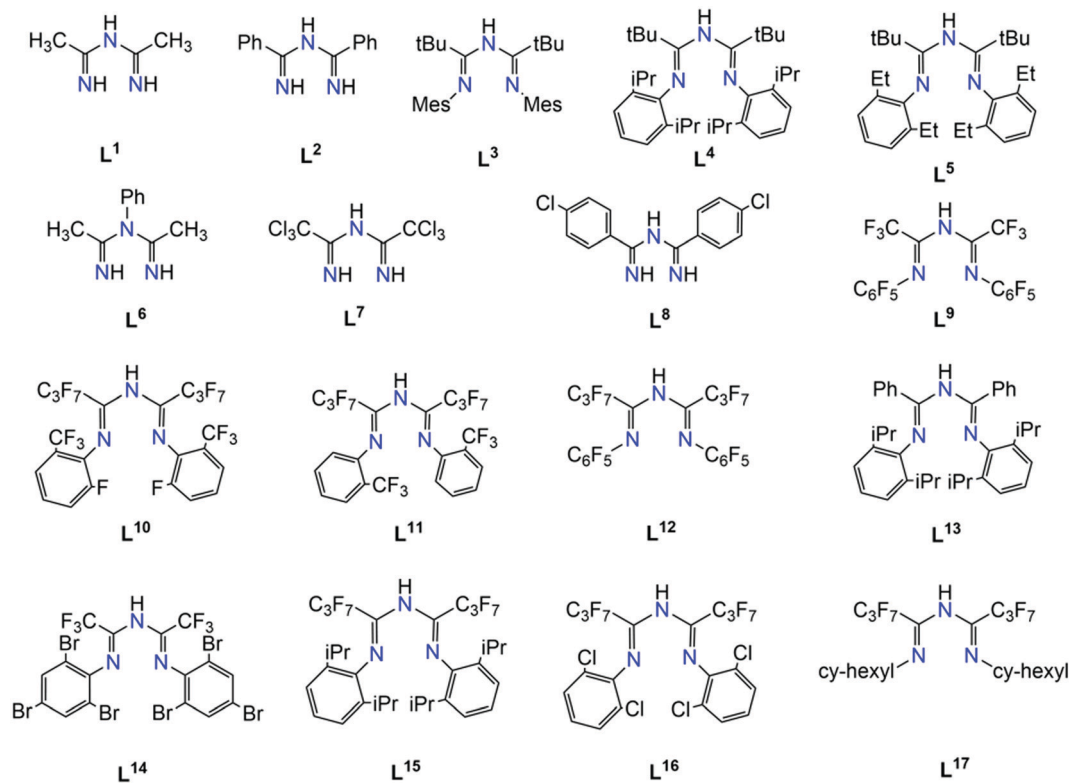
Replacement of hydrogen as the R_2 substituent by bulkier aryl or aromatic groups can lead to structurally interesting **ImAm** ligands such as L^3 , L^4 and L^5 . Pernik *et al.* employed these bulky frameworks in the formation of mononuclear and dinuclear complexes of Li^{I} , K^{I} , Ca^{II} and Mg^{II} .⁴⁰ In contrast to the previous examples, in this work the ligands were pre-synthesised prior to reactions with metal salts. In all these complexes, the deprotonated ligands act in the U-shaped coordination mode. In a similar manner, the H-atom of the R_3 group can also be replaced, as demonstrated by Bolaño *et al.*⁴¹ They reported the *in situ* synthesis of the L^6 ligand upon coordination to a Rh^{III} center leading to the formation of the trication $[(\text{tbpy})_2\text{Rh}^{\text{III}}(\text{L}^6)][\text{OTf}]_3$ (**2**) (tbpy = 4,4'-di-*tert*-butyl-2,2'-bipyridine; OTf = CF_3SO_3^- ; Fig. 2), representing one of the few examples of **ImAm**-based complexes where the central N-atom is substituted. In **2**, the Rh^{III} metal center has a slightly distorted octahedral geometry with two tbpy and one L^6 ligand completing the coordination sphere.

Replacement of the H-atoms in the methyl groups of L^1 with chlorides affords the ligand L^7 , which was also prepared *in situ* as reported by Shixaliyev *et al.* where they isolated a family of mononuclear square planar complexes of the generalized type $\{\text{M}^{\text{II}}[\text{L}^7]_2\} \cdot 2(\text{Me}_2\text{SO})$, where $\text{M} = \text{Ni}^{\text{II}}$ (**3**), Cu^{II} (**4**) or Pd^{II} (**5**) (Fig. 3).⁴² Strong H-bonds of the form $\text{N}-\text{H} \cdots \text{O}$ (where O belongs to a dimethylsulfoxide (Me_2SO) lattice solvent molecule) as well as $\text{Cl} \cdots \text{Cl}$ halogen bonds create an elaborate 2D supramolecular network. Due to the chemical and thermal stability as well as enhanced solubility in halocarbons (*i.e.*, dichloromethane, chloroform, chlorobenzene, *etc.*) that these halogenated **ImAm** complexes exhibit, the researchers explored the catalytic properties of similar mononuclear complexes of L^7 of the general formula $[\text{M}^{\text{III}}(\text{L}^7)_3] \cdot x(\text{Me}_2\text{SO})$, where $\text{M} = \text{Mn}^{\text{III}}$ (**6**) with $x = 1$; Fe^{III} (**7**) with $x = 2$; and Co^{III} (**8**) with $x = 2$.⁴³ As in the previous study, **6-8** were prepared using a one pot synthetic procedure. In each complex, three L^7 ligands are coordinated to the metal center in a bidentate fashion leading to a slightly distorted octahedral geometry about the metal ions. As previously observed, the H-atoms of the amine groups were engaged in strong H-bonding with a Me_2SO lattice solvent molecule and halogen bonds lead to the formation of a supramolecular organization. These complexes were tested as catalyst precursors for the additive-free microwave (MW) assisted homogeneous oxidation of secondary alcohols (such as 1-phenylethanol) to the corresponding ketones by aqueous *tert*-butylhydroperoxide (TBHP) or H_2O_2 without an added solvent. Although the Mn^{III} and Fe^{III} complexes showed notable catalytic properties, the best yields were achieved for the Co^{III} analogue.

The introduction of bulkier groups such as chlorobenzene (*i.e.*, ligand L^8) in the R_1 groups was achieved by Kopylovich *et al.* in a one pot synthesis starting with halogenated benzoni-triles in the presence of PdCl_2 .⁴⁴ This led to the formation of the cationic $[\text{Pd}^{\text{II}}(\text{L}^8)_2]^{2+}$ (with the ligand L^8 being in the neutral form) and the neutral square planar Pd^{II} complex $[\text{Pd}^{\text{II}}(\text{L}^8)_2]$ (**9**, where both ligands are in the monoanionic form).

The presence of halogen atoms in both the R_1 and R_2 groups of **ImAm** has inspired the synthesis of ligands with rich halogen





Scheme 9 Examples of **ImAm** ligands that do not possess additional donor atoms in the R_1 , R_2 or R_3 groups.

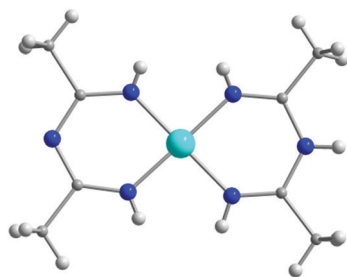


Fig. 1 The molecular structure of the cationic complex **1**,^{39b} where the L^1 ligand exists in both its neutral and anionic form. Lattice solvent molecules and counter ions have been omitted for clarity. Colour code: C: grey, H: white, N: blue, Ni^{II}: light blue. Reproduced using reported cif.

content due to their potential catalytic applications.^{42–45} To that end, ligands such as L^9 , L^{10} , L^{11} and L^{12} have been prepared and, upon reaction with Cu^I sources, have led to the isolation of mononuclear complexes. In the work of Ponduru *et al.*, $[Cu^I(L^9)MeCN]$ (**10**) was isolated upon treatment of L^9 with CuO in acetonitrile.^{45a} In **10**, the ligand acts in a chelating fashion in the U-shape coordination mode and the copper center adopts a trigonal planar geometry (Fig. 4). This complex was tested as a possible catalyst for the oxidation of anthracene, naphthalene and pyrene to the corresponding quinones by usage of H_2O_2 as an oxidant under mild conditions without the presence of an acid co-catalyst. Similar work was also published by the same group utilizing L^{10} , L^{11} and L^{15} as

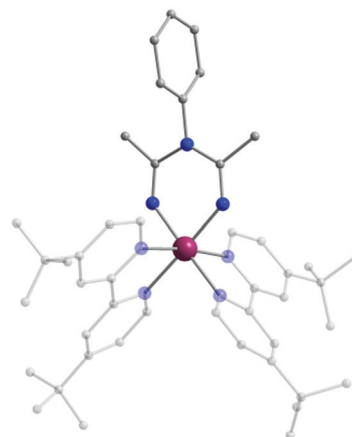


Fig. 2 The molecular structure of the cationic complex **2**.⁴¹ For clarity reasons H-atoms, disorder conformers and lattice counter ions have been omitted, as well as partial transparency has been employed. Colour code: C: grey, N: blue, Rh^{III}: purple. Reproduced using reported cif.

ligands, which led to the isolation of analogous mononuclear complexes of Cu^I .^{45b}

Despite the prominence of U-shape chelating coordination mode that these ligands exhibit, other coordination modes have also been reported. In the work of Dias *et al.* a series of mononuclear Cu^I and Ag^I complexes were isolated.⁴⁶ In $[Ag^I(L^9)(CN^tBu)_2]$ (**11**), the L^9 ligand coordinates in a terminal W-shaped fashion through the central N-atom with the three-coordinate Ag metal



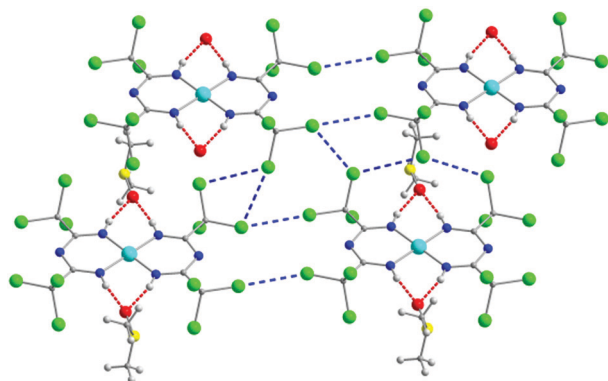


Fig. 3 The 2D supramolecular organization of **3**.⁴² Cl...Cl interactions are represented as dashed blue lines, while H-bonds as dashed red lines. Colour code: C: grey, N: blue, O: red, Cl: green, Ni^{II}: light blue. Reproduced using reported cif.

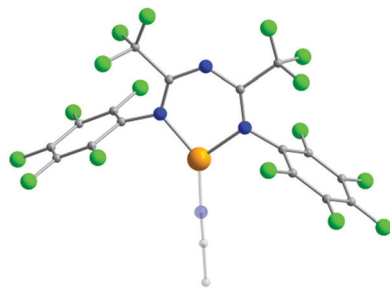


Fig. 4 The molecular structure of the cationic complex **10**.^{45a} For clarity reasons H-atoms and disorder conformers have been omitted, as well as partial transparency has been employed. Colour code: C: grey, N: blue, F: green, Cu^I: orange. Reproduced using reported cif.

center adopting a distorted trigonal planar geometry (Fig. 5a). Changing from L^9 to L^{12} , which possesses bulkier R_1 substituents, leads to the formation of the isostructural Ag^I complex $[Ag^I(L^{12})(CN^tBu)_2]$ (**12**). By addition of triphenylphosphine to Ag_2O in the presence of L^9 and L^{12} , two isostructural complexes of Ag^I $[Ag^I(L^9)(PPh_3)_2]$ (**13**) and $[Ag^I(L^{12})(PPh_3)_2]$ (**14**) were isolated. In these complexes both ligands act in a terminal fashion through one of the terminal N-atoms, while the three-coordinate Ag^I metal center is in a distorted trigonal planar geometry (Fig. 5b). Apart from these, they also reported a mononuclear Cu^I complex, $[Cu^I(L^9)(CN^tBu)_2]$ (**15**) with L^9 , where the ligand binds to the metal in a U-shape manner as previously illustrated. In **15**, the metal center is four-coordinate in a slightly distorted tetrahedral geometry. This work highlights the versatility of **ImAm** ligands, which can bind to metal ions in several different manners. Additionally, computational analysis at the PM3 level of theory on the U- and W-shape coordination motifs has shown that these are the lowest energy conformations.⁴⁶ Moreover, due to the almost evenly distributed electron density among the three N-atoms within the HOMO, these ligands are flexible enough to adapt relatively easily by altering their coordination mode in order to placate any steric-electronic demands of the coordination fragment.

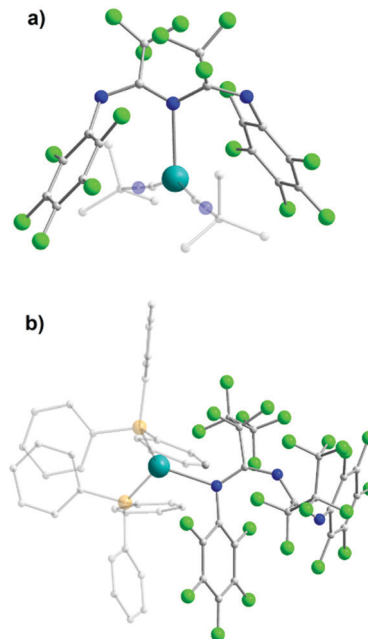
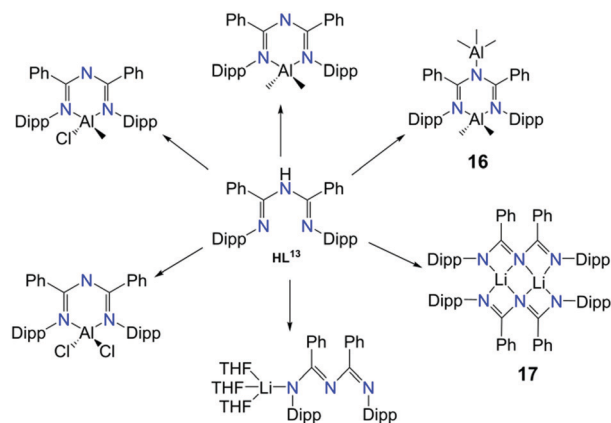


Fig. 5 The molecular structure of the cationic complex **11** (a) and **14** (b).⁴⁶ For clarity reasons H-atoms and disorder conformers molecules have been omitted, as well as partial transparency has been employed. Colour code: C: grey, N: blue, F: green, Ag^I: light blue. Reproduced using reported cifs.

The coordination possibilities of **ImAm** ligands that bear non coordinating groups was fully explored by Bakthavachalam *et al.* in a series of Li^I and Al^{III} complexes with L^{13} .⁴⁷ They illustrated that by employing various reaction conditions, mono- and dinuclear complexes of Li^I and Al^{III} metal centers could be isolated (Scheme 10). Due to structural similarities only three of the complexes are discussed in detail. In complex $[L^{13}AlMe_2(AlMe_3)]$ (**16**) the L^{13} ligand acts as a *syn-anti* tridentate bridging moiety by utilizing all its donor atoms. This is the only example where the **ImAm** ligand acts as a *syn-anti* bridge when the R_1 groups do not possess donor atoms. Additionally, by using Li^I instead of Al^{III} , the authors were able to isolate the dinuclear Li^I complex $[(L^{13}Li)_2]$ (**17**) from Et_2O . In **17**, each Li^I



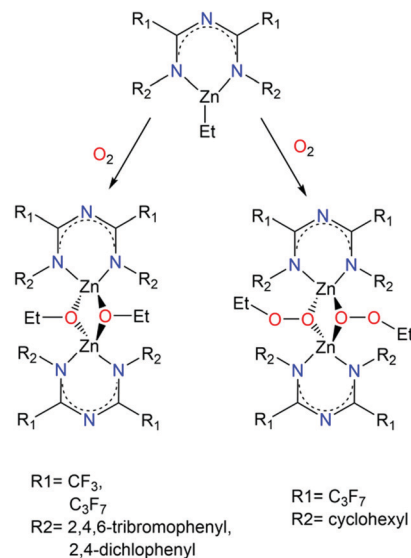
Scheme 10 The synthesis of mononuclear and dinuclear complexes of Al^{III} and Li^I with HL^{13} .⁴⁷



ion is bound to four N-atoms in a highly distorted tetrahedral geometry with both ligands adopting a *syn-syn* W-shape coordination motif, which stabilizes the two Li^{I} metal centers. In contrast, using THF instead of Et_2O a mononuclear Li^{I} complex has been obtained, $[\text{L}^{13}\text{Li}(\text{THF})_3]$, where L^{13} coordinated in the terminal fashion through only one of the N atoms.

In an effort to explore the coordination chemistry of HL^{13} with 3d transition metal ions, Nareddula and co-workers focused their attention on the isolation of Zn^{II} complexes.^{48a} By reacting HL^{13} with ZnEt_2 they were able to isolate the mononuclear complex $[\text{L}^{13}\text{ZnEt}]$ (**18**). Here, L^{13} is deprotonated and acts as a chelating (U-shaped) ligand, stabilizing the Zn^{II} ion in a slightly distorted trigonal planar geometry. Interestingly, by changing the Zn^{II} metal source to $\text{Zn}(\text{OAc})_2$ the authors were able to isolate a tetranuclear Zn^{II} cluster $[(\text{L}^{13})_2\text{Zn}_4\text{O}(\text{OAc})_4]$ (**19**) (Fig. 6). In this complex, the four Zn^{II} centers are connected *via* a central $\mu_4\text{-O}^{2-}$ bridge. Two bridging acetato ligands also contribute to the coordination sphere of each metal center, while the L^{13} coordinates in a *syn-syn* W-shape manner, bridging two Zn^{II} metal centers. This coordination motif suggests that **ImAm** ligands can also act as bridging moieties for the formation of high nuclearity complexes in 3d metal chemistry. Additionally, these complexes were also tested as efficient catalysts for the ring opening polymerization (ROP) of ϵ -caprolactone.^{48a}

A series of mononuclear and dinuclear complexes of Zn^{II} was reported by Kulkarni *et al.* where reaction of L^{14} , L^{15} , L^{16} and L^{17} with ZnEt_2 led to the isolation of a mononuclear coordination complex of the generalized formula $[\text{L}^x\text{ZnEt}]$ (where $x = 14, 15, 16$ or 17).^{48b} While these complexes are similar to **18**, upon exposure to O_2 for a limited time followed by recrystallization at specific temperatures, dimerization of the complexes containing L^{14} , L^{16} and L^{17} occurred (Scheme 11). In particular, when L^{14} and L^{16} ligands were employed, dimerization was achieved through the formation of ethoxy bridges, while in



Scheme 11 The synthesis of Zn^{II} ethoxide and ethylperoxide dimers.^{48b}

the case of L^{17} the formation of ethylperoxy bridges was observed. In case of L^{15} , when the corresponding complex was exposed to O_2 , a mixture of intractable products was obtained. This difference can be attributed to the nature of the R_2 groups that bind to the terminal N-atoms of the **ImAm** framework. For example, the L^{17} ligand bears electron-donating groups making it more capable of stabilizing the zinc-ethylperoxy complex. In contrast, the halogen rich substituents in L^{14} and L^{16} are electron-withdrawing and therefore lead to the formation of the zinc-ethoxy complexes.

Overall, the coordination chemistry of these ligands has not been fully explored. This may be attributed to the absence of donor atoms in the R_1 and R_2 groups and the energetically favored U-shape coordination mode, leaving little room for isolation of metal complexes with higher nuclearities. Nonetheless, polynuclearity can be achieved by the inclusion of donor atoms in the R_1 and R_2 moieties.

3.2 ImAm chelates containing additional donor atoms in the R_1 positions

Apart from the three N-atoms of the N_3C_2 moiety, efforts to synthesize polydentate imidoamidine ligands have led to impressive results. Of particular relevance to the current review, the ligands presented in Scheme 12 will be discussed in turn all of which possess multiple sites for coordination.

Due to the presence of two pyridyl substituents in L^{18} , both bidentate and tridentate coordination sites are created. Coordination in a bidentate fashion has been well known since several groups managed to isolate complexes by the *in situ* synthesis of L^{18} *via* metal-assisted transformations.⁴⁹ Such a preparative manner that is reliant on the serendipitous assembly of the ligand framework does not enable full exploration of the capabilities of L^{18} . On the other hand, having the ligand framework pre-made and then utilized in coordination chemistry provides better control over the synthetic process

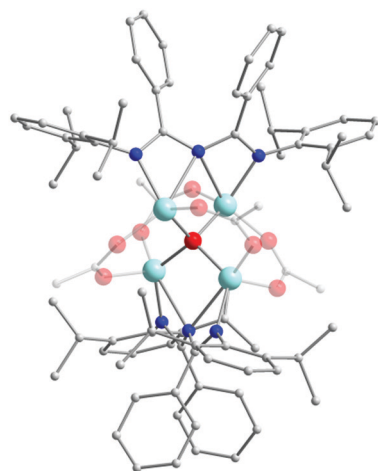
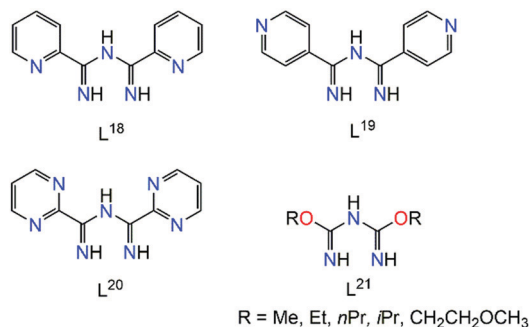


Fig. 6 The molecular structure of the cationic complex **19**.^{48a} For clarity reasons H-atoms and disorder conformers have been omitted, as well as partial transparency has been employed. Colour code: C: grey, N: blue, O: red Zn^{II} : pale blue. Reproduced using reported cif.





Scheme 12 Examples of polydentate **ImAm** ligands with additional coordination groups in the R_1 positions.

and overall yields. The importance of this was highlighted in the development of a series of mononuclear coordination complexes with Mn^{III} , Fe^{III} and Co^{III} under various synthetic conditions, which we recently reported.⁵⁰ In particular, when reactions were carried out under basic or neutral conditions mononuclear complexes with a generalized molecular formula $[M^{III}(L^{18})_3]$ (where $M = Mn^{III}$ (**20**), Fe^{III} (**21**) or Co^{III} (**22**)) (Fig. 7a) were isolated. In these complexes, the monoanionic ligand binds through its bidentate chelate in a U-shape fashion, stabilizing the trivalent metal ions in an octahedral molecular geometry. Conversely, under acidic conditions the tridentate coordination pocket of L^{18} could be accessed, leading to the isolation of the complexes $[M^{II}(L^{18})Cl_2]$ (where $M = Mn^{II}$ (**23**) or Co^{II} (**24**)) and $[Fe^{III}(L^{18})Cl_3]$ (**25**) in which the ligand was not deprotonated. The additional hydrogen atom on L^{18} in **23–25** is shared between the imino groups (50% occupancy due to the presence of a crystallographic C_2 axis) effectively blocking the bidentate pocket, thus leading to coordination to the metal ions through the two pyridyl and the central amidine nitrogen atoms. In the Mn^{II} (**23**) and Co^{II} (**24**) analogues metal oxidation was not observed, as was the case under basic conditions, leading to metal ions in which bivalent oxidation states are preserved. Two chloro terminal ligands complete their coordination spheres (Fig. 7b). In the case of the iron derivative **25**, due to the trivalent iron starting material, Fe^{III} has an octahedral molecular geometry with three chloro terminal groups completing its coordination sphere (Fig. 7c). Interestingly, neither the change of the transition metal nor the solvent seemed to affect the products isolated, although a change of

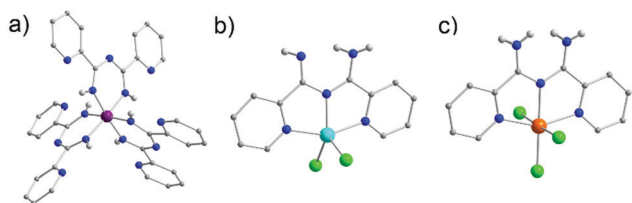
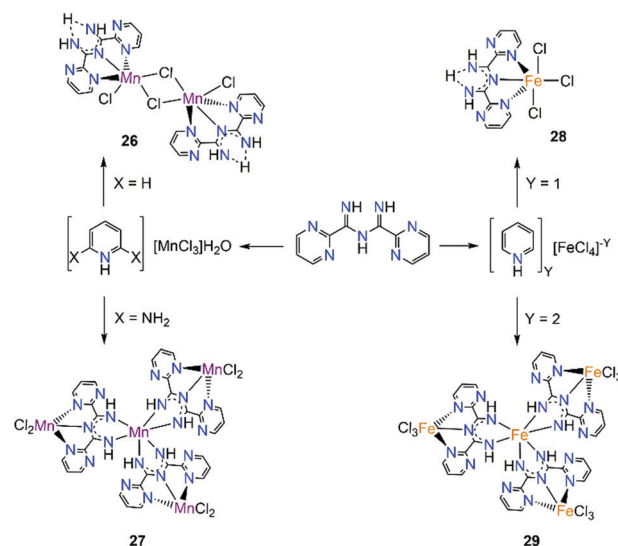


Fig. 7 The molecular structure of the complex **20** (a), **24** (b) and **25** (c).⁵⁰ For clarity reasons H-atoms and lattice solvent molecules have been partially omitted. Colour code: C: grey, N: blue, Cl: green, H: light grey, Mn^{III} : purple, Co^{II} : teal, Fe^{III} : orange. Reproduced using reported cifs.

the latter from MeCN to dichloromethane (DCM) affected the crystal packing.

Intrigued by these findings, we sought to isolate polynuclear coordination complexes with similar ligand frameworks, such as L^{20} where the pyridyl rings have been replaced by pyrimidyl groups.⁵¹ In this work, two coordination complexes of $Mn^{II/III}$ and two coordination complexes of Fe^{III} were isolated and their magnetic behaviour along with the magnetic behaviours of the previously isolated **20** and **21**, were investigated (Scheme 13). The first complex isolated, $[Mn_2^{II}(\mu-Cl)_2(L^{20})_2Cl_2]$ (**26**), is a centrosymmetric dinuclear complex with two $\mu-Cl^-$ the two Mn^{II} centers. The neutral ligand coordinates in a similar fashion as in the case of **23**, through its tridentate binding pocket, while the slightly distorted octahedral coordination geometry of the Mn^{II} centers is completed by a terminal chloride. By changing to less acidic conditions, the tetranuclear complex $[Mn_4^{III}Mn_3^{II}(L^{20})_3Cl_6]$ (**27**) was isolated. In this complex, three deprotonated L^{20} ligands are bound to the central Mn^{III} ion in a bidentate chelate fashion as in the case of complex **20**. Each ligand is simultaneously coordinated through its tridentate pocket to another Mn^{II} metal center while two terminal chloro co-ligands complete its coordination sphere. By employing similar reaction conditions as in the case of **25** with Fe^{III} , the mononuclear $[Fe^{III}(L^{20})Cl_3]$ (**28**) was obtained. This shares structural resemblance with complex **25**, since the neutral ligand is coordinating again in a mer-fashion, with three Cl^- ions completing the coordination sphere of Fe^{III} . The tetranuclear Fe^{III} analogue of **27**, $[Fe_4^{III}(L^{20})_3Cl_9]$ (**29**), was achieved by changing the reaction solvent as well as the starting material of iron. Complexes **27** and **29** share structural similarities such as the coordination fashion of the deprotonated L^{20} and the coordination environment of the central metal ion. However, in the case of $[Fe_4^{III}(L^{20})_3Cl_9]$ (**29**), all four iron ions are in their 3+ oxidation state and in order to satisfy charge balance considerations each of the peripheral Fe^{III} ions is bound to three Cl^- terminal ligands.



Scheme 13 Synthesis of complexes **26**, **27**, **28** and **29** employing L^{20} .⁵¹



Direct current (dc) magnetic susceptibility measurements performed on complexes **20**, **21**, **26**, **27**, **28** and **29** (Fig. 8) as well as Mössbauer spectroscopy performed on **29** (Fig. 9), revealed that these **ImAm** ligands can stabilize both high and low spin configurations of these metal centers. For instance, in the case of **28** and **29** the central metal ions have a low-spin configuration due to their N-rich coordination environment provided by the **ImAm** ligands. On the contrary, the peripheral metal ions that also bind to Cl⁻ adopt high-spin configurations. The incorporation of such **ImAm** frameworks that can promote both ferro- and antiferromagnetic interactions between the metal centers can prove to be most useful in the field of Molecular Magnetism, as such examples are rare. Therefore, synthetic strategies as those implemented in these cases can prove helpful in the targeted design of polynuclear species with different spin states configurations.

Although these synthetic procedures have yielded impressive results, we wanted to explore the potential of the aforementioned ligands under inert conditions. Recently, both L¹⁸ and L²⁰ were employed to isolate mononuclear as well as polynuclear Co^{II} complexes, respectively.⁵² Under basic reaction conditions in an inert environment mononuclear [Co^{II}(L¹⁸)₂] (**30**) was obtained (Fig. 10a). In contrast to complex **22**, the Co^{II} has an almost ideal square planar geometry since both ligands coordinate in a bidentate fashion and the metal ion is in its 2+

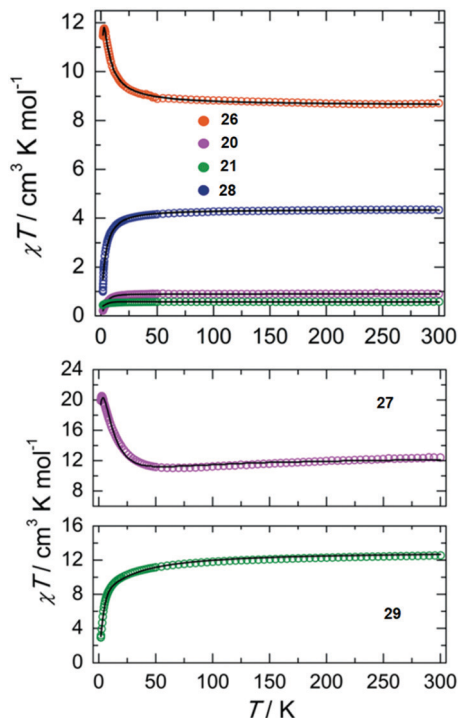


Fig. 8 Top: Temperature dependence of the χT product of **20** (purple), **21** (green), **26** (orange) and **28** (blue) with χ being the molar magnetic susceptibility per molecule as defined by M/H . Bottom: Temperature dependence of the χT product of **27** (purple) and **29** (green) with χ being the molar magnetic susceptibility per molecule as defined by M/H . In both plots the black solid lines represent the best-fits to the data. Adapted with permission from ref. 51. Copyright 2019 Wiley.

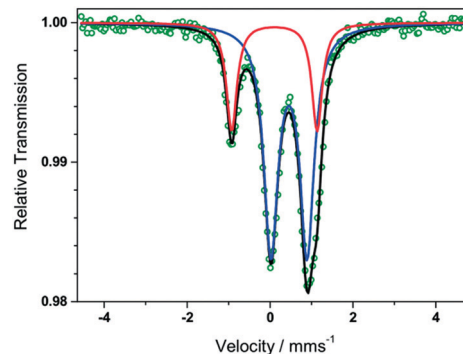


Fig. 9 Zero-field Mössbauer spectrum of **29** recorded at 80 K; the collected experimental data is indicated as green circles. The blue and red lines correspond to the fit of outer high spin and the central low spin Fe^{III} ions, respectively, showing that there are two different species of Fe^{III} in the complex. The overall sum of the Fe^{III} ions is represented by the black line. Adapted with permission from ref. 51. Copyright 2019 Wiley.

oxidation state, as confirmed by both charge balance considerations and dc magnetic susceptibility measurements. Replacement of L¹⁸ with L²⁰, led to the isolation of [Co^{II}(L²⁰)₂Cl₄] (**31**) and [Co^{II}(μ -Cl)₂(L²⁰)₄Cl₆] (**32**) (Fig. 10b and c, respectively). Complex **31** is a linear trinuclear Co^{II} complex, where both the bidentate and tridentate coordination pockets of each deprotonated L²⁰ moiety is occupied by a Co^{II} ion. The central Co^{II} adopts a square planar geometry identical to that of **30**. The two peripheral Co^{II} ions have geometries similar to the one observed for complex **24**, with three N-donor atoms of the L²⁰ ligand and the two terminal Cl⁻ atoms completing their coordination sphere. Interestingly, the degree of hydration of the reaction solvent (MeOH) is the key difference between isolating complex **31** and **32**. Complex **32** can be described as two trinuclear moieties (similar to **31**) bridged by two μ -Cl⁻. The asymmetric unit of **32** consists of three unique Co^{II} centers. Within this unit, the two L²⁰ ligands coordinate in the same manner as in the case of **31**. However, the central Co^{II} is pentacoordinate due to the bridging Cl⁻, resulting in a slightly distorted square pyramidal geometry.

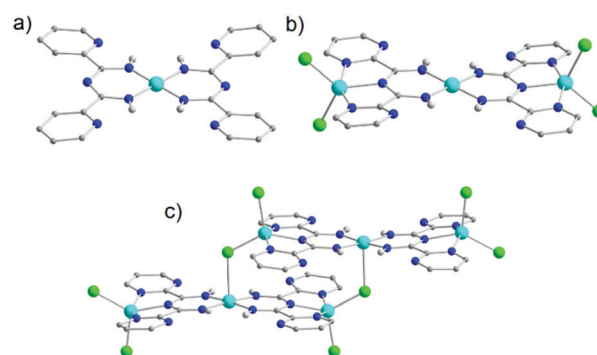


Fig. 10 The molecular structures of complexes **30** (a), **31** (b) and **32** (c).⁵² For clarity reasons H-atoms and lattice solvent molecules have been omitted. Colour code: C: grey, N: blue, Cl: green, Co^{II}: tale. Reproduced using reported cifs.



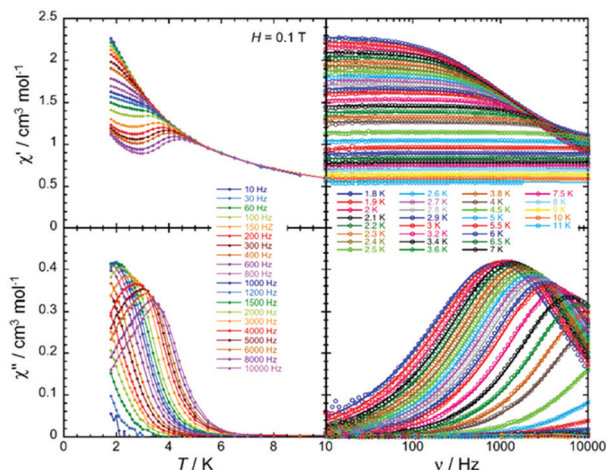


Fig. 11 Temperature (left) and ac frequency (right) dependences of the in-phase (χ' , top) and out-of-phase (χ'' , bottom) ac susceptibility in the temperature region of 1.85 to 10 K and between 10 and 10 000 Hz for **32** in a 0.1 T dc field. Solid lines serve, on the left plots as guides to the eye while on the right they show the generalized Debye fit of the ac data. Adapted with permission from ref. 52. Copyright 2020 Royal Society of Chemistry.

Magnetic susceptibility measurements as well as magnetostructural correlations of the $\{\text{Co}_3^{\text{II}}\}$ aggregates in **31** and **32** reveal a unique combination of low spin Co^{II} (center) and high spin Co^{II} (peripheral) configurations. Additionally, complex **32** exhibits SMM behavior (Fig. 11) under a small applied field, where Raman, quantum tunneling of the magnetization and direct process are responsible for the slow relaxation of the magnetization (Fig. 12).

Recently, Piers and co-workers presented two mononuclear Ni^{II} complexes with L^{18} .⁵³ The first was the neutral complex $[\text{Ni}^{\text{II}}(\text{L}^{18})_2]$ (**33**) in which the monoanionic bidentate chelate ligands stabilize the square planar Ni^{II} ion (Fig. 13a). The one e^- reduction of **33** results in the formation of $[\text{K}(\text{crypt})][\text{Ni}^{\text{II}}(\text{L}^{18})_2]$ (**34**) in which the **ImAm** ligand is in its reduced form. As supported by density functional theory (DFT) calculations and electron paramagnetic resonance (EPR) spectra the radical species is delocalized almost exclusively onto the ligand framework (Fig. 13b and c, respectively). Controlled potential electrolysis revealed that

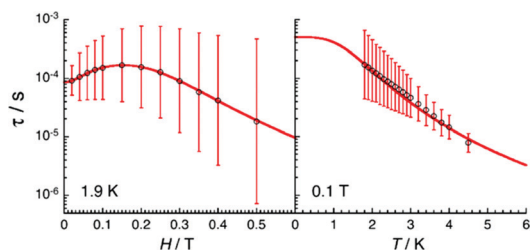


Fig. 12 Dc-field (left) and temperature (right) dependences of the relaxation time for **32** estimated from the generalized Debye fits of the ac data shown in Fig. 10 collected at 1.9 K and under an 0.1 T applied field, respectively. The vertical solid bars are the estimated standard deviations of the relaxation time. The solid red line represents the best fit. Adapted with permission from ref. 52. Copyright 2020 Royal Society of Chemistry.

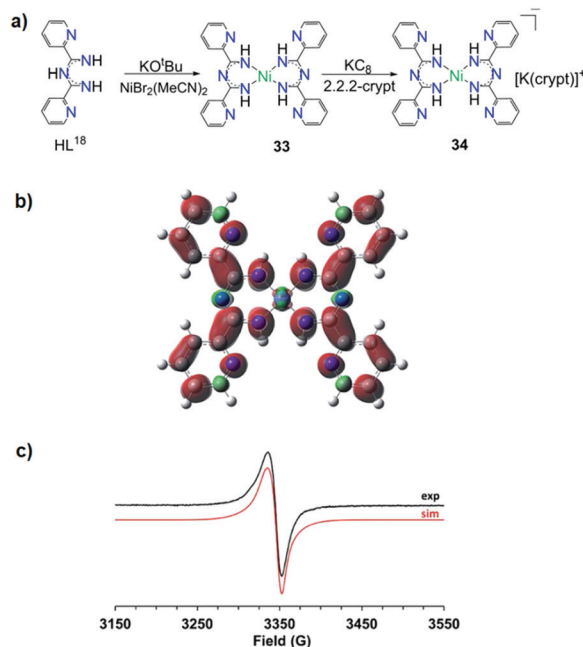


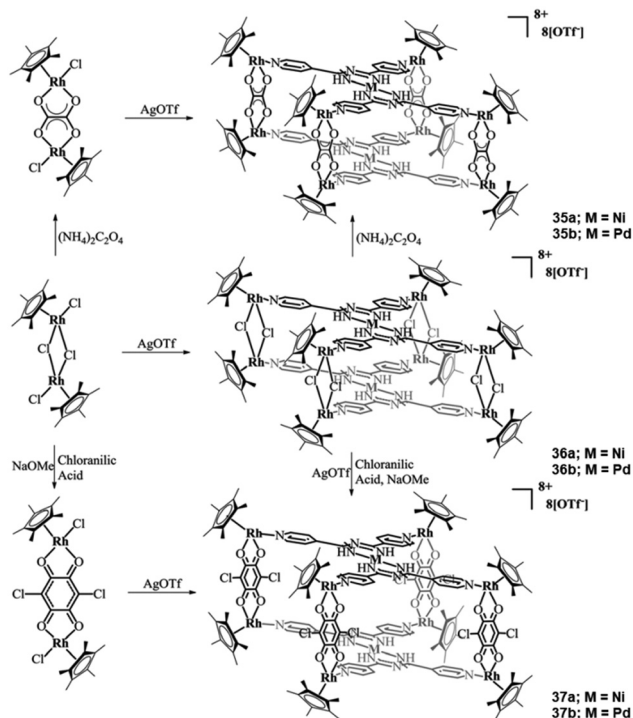
Fig. 13 (a) The synthetic procedure that led to the isolation of complexes **33** and **34**; (b) the Mulliken spin density plot of **34**; (c) experimental (back) and simulated (red) X-band EPR spectra of **34**. Reprinted with permission from ref. 53. Copyright 2019 Royal Society of Chemistry.

reduced **33** can convert CO_2 to CO , highlighting that the redox activity of this ligand framework can play a significant role in the reduction of CO_2 *via* the utilization of 3d metal complexes.

By changing the pyridine N-atom from the *ortho*- to the *para*-position, L^{19} can be formed. Jin and co-workers reported a series of heterometallic cationic complexes of the generalized formula of $\{[\text{Cp}^*\text{Rh}_2\text{X}_2]_4[\text{M}^{\text{II}}(\text{L}^{19})_2]_2\}$ (**35–37**) (where $\text{X} = \text{Cl}$, $(\mu\text{-}\eta^4\text{-C}_2\text{O}_4)$, $(\mu\text{-chloroanilic acid})$ and $\text{M} = \text{Ni}^{\text{II}}$ or Pd^{II}).⁵⁴ The authors isolated six heterometallic complexes (**3d/4d** and **4d/4d**) following the generalized synthetic procedure outlined in Scheme 14. In all six cationic complexes, the L^{19} ligands coordinate in the same fashion, the only difference being the bridging co-ligand between the Rh metal centers or the exchange of Ni^{II} with Pd^{II} . Thus, for the purposes of this review article, only **36a** will be discussed in detail. In this complex, each L^{19} ligand coordinates to the Ni^{II} metal centers through its bidentate U-shape site, stabilizing the Ni^{II} ions in a square planar geometry. Four coplanar *N*-pyridyl atoms in L^{19} coordinate with four $\{\text{Cp}^*\text{Rh}\}$ moieties, with each of the Cp^*Rh moieties bridged through two chloride co-ligands with a second Cp^*Rh moiety that has been stabilized in the “corners” of the L^{19} ligand. As a result, two $\text{Ni}(\text{L}^{19})_2$ moieties stack in a parallel fashion, forming a double decker bridged by the $\{\text{Cp}^*\text{Rh}_2\text{Cl}_2\}$.

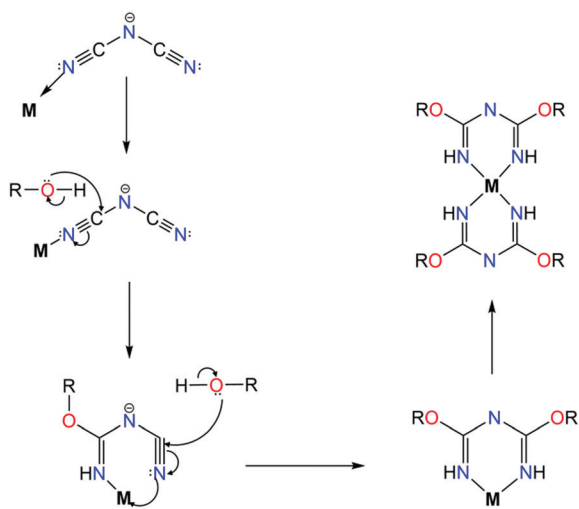
As evident from the aforementioned complexes, the introduction of functional groups with N-atoms has yielded impressive results based on stepwise inorganic syntheses. Attempts to incorporate other functional groups such as RO^- , have led to the formation of a small family of mononuclear complexes. The main difference in this category of ligands, such as L^{21} , is that these chelates are synthesized *in situ*, *via* a metal assisted





Scheme 14 The synthetic routes that led to the isolation of the organometallic heterometallic complexes, where $M = \text{Ni}^{\text{II}}$ or Pd^{II} . Adapted with permission from ref. 54. Copyright 2017 Royal Society of Chemistry.

pathway. In particular, the synthesis of these 2,4-alkoxy-1,3,5-triazapentadienes is based on the reaction of sodium dicyanamide with alcohols in the presence of a transition metal source, leading to mononuclear complexes of Cu^{II} ^{55a-c} or Ni^{II} ^{55d} with a generalized formula $[\text{M}^{\text{II}}(\text{L}^{21})_2]$ (where $M = \text{Cu}^{\text{II}}$ or Ni^{II}). Efforts to understand the mechanisms behind this ligand formation, such as the one given in Scheme 15, have been made by several groups.^{55a,b} In these complexes, the ligand acts in the chelate,



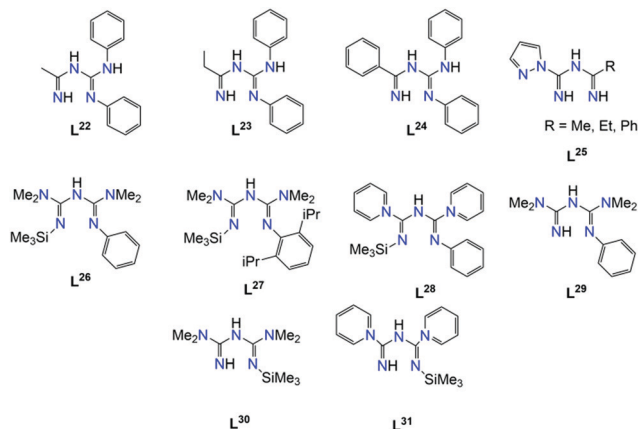
Scheme 15 The proposed mechanism for the *in situ* synthesis of 2,4-alkoxy-1,3,5-triazapentadiene ligands, where $M = \text{Cu}^{\text{II}}$ or Ni^{II} and $R = \text{Me}$, Et , $n\text{Pr}$, $i\text{Pr}$ or $\text{CH}_2\text{CH}_2\text{OCH}_3$.^{55b}

U-shape, coordination mode with the divalent transition metal ion adopting a square planar geometry. In the case of Ni^{II} , by employing pyrazole-based co-ligands, dinuclear and trinuclear Ni^{II} complexes were isolated with the ligand coordinating in the same fashion. The synthesis of these ligands has only been achieved *in situ* and efforts to utilize the alkoxy groups as possible coordination sites remain an ongoing challenge given there is no control over this synthetic approach.

ImAm ligands can provide promising complexes for catalytic applications like their β -diketone analogues.^{55c} For this reason, some of the 2,4-alkoxy-1,3,5-triazapentadienato Cu^{II} complexes have been tested as efficient catalysts for the nitroxyl radical mediated mild aerobic oxidation of alcohols to the corresponding aldehydes and for the solvent-free microwave-assisted synthesis of ketones from secondary alcohols with *tert*-butylhydroperoxide as oxidant.^{55a} In addition to that, other **ImAm** based complexes have been studied as catalysts for the mild oxidation of alkanes with H_2O_2 as an oxidant, pyridine as promoting agent and cyclohexane as a main model substrate.^{55c}

4. The coordination chemistry of asymmetrical imidoyl amidine chelates

As evident from the aforementioned examples, the presence of different substituents plays a key role in the coordination chemistry of **ImAm** based ligands. By introducing different R groups, desired structural motifs can be targeted. For instance, such frameworks can have an R_1 group, such as pyridine, that stabilize metal ions through chelating coordination and a different R_1' substituent, such as hydroxyl, that can bridge multiple metal centers. Moreover, apart from an additional coordination site, the central N-atom can also take part in facile acid/base equilibria, providing pH sensing properties to complexes.⁵⁶ This can prove to be important for future application of such complexes that exhibit luminescence properties. In Scheme 16 the ligands that will be discussed in this section are presented.

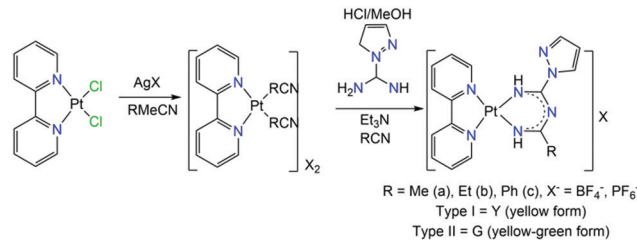


Scheme 16 The asymmetrical **ImAm** chelates that will be discussed in this section.



In an effort to explore the luminescent properties of asymmetric **ImAm** ligands with Pt^{II}, Eliseev *et al.* synthesized a series of mononuclear Pt^{II} complexes [Pt^{II}(dpa)(L^x)](OTf) (**38**) (where $x = 22, 23, 24$; dpa = 2,2'-dipyridylamine; OTf = SO₃CF₃).⁵⁶ Due to the fact that in all three cationic complexes the ligands act in the same coordination mode and since all three share structural similarities, only **38a** will be discussed with L²² as ligand. As shown in Fig. 14 (top), the L²² ligand, which was synthesized *in situ*, is in its monoanionic form, coordinating in the chelate, U-shape, coordination fashion. Four N-atoms (two from the **ImAm** ligand and two from the dpa co-ligand) complete the coordination sphere of the Pt^{II} ion leading to a slightly distorted square planar molecular geometry. Surprisingly, these complexes showed intense emission in the solid state (Fig. 14, bottom), but they were non-emissive in solution. This can be explained by the square planar configuration of Pt^{II} in which an unoccupied d orbital is located perpendicular to the plane. Theoretically, the empty orbital can be subject to coordination by electron donors of the lattice solvent molecules, which leads to dominant non-radiative deactivation pathways.

Since the emissive states of Pt^{II} complexes can be affected dramatically by the bulkiness of the substituents on the **ImAm** ligands, Tanaka *et al.* synthesized a series of Pt^{II} based complexes with L²⁵.⁵⁷ The introduction of the pyrazole substituent



Scheme 17 Synthesis of [Pt^{II}(bpy)(L^{25R})]X (**39R**) (where bpy = 2,2'-bipyrimidine, R = Me (a), Et (b) or Ph (c) and X = BF₄⁻, PF₆⁻).⁵⁷

is key, for the careful tuning of the basicity of pyrazolate ligands and enables control over the structure of mononuclear Pt^{II} complexes. The general synthetic procedure that led to the isolation of these complexes is given in Scheme 17. The ligand is generated *in situ*, in a nitrile solution with a pyrazole derivative bearing the amidine group in the presence of the metal salt. Formation of this ligand is achieved by the coupling of the coordinated nitriles with the pyrazolecarboxamidine. The final products of these reactions, depending on the nitrile in use, have a generalized formula of [Pt^{II}(bpy)(L^{25R})]X (**39R**) (where bpy = 2,2'-bipyrimidine, R = Me (a), Et (b) or Ph (c) and X = BF₄⁻, PF₆⁻). Different recrystallization methods led to two different crystal phases; type I, which exhibits yellow luminescence and type II, which exhibits yellow-green luminescence. In Fig. 15 (top), the solid state emission spectra of these complexes is given.

Based on time-dependent density functional theory (TD-DFT) calculations, it was found that the lowest energy absorption bands of these complexes can be assigned to a combination of ligand-to-ligand charge-transfer (LLCT) and metal-to-ligand charge transfer (MLCT) transitions. These findings highlight the prosperity of aromatic **ImAm** ligands in the development of complexes for optical applications. Interestingly, these complexes also exhibited unique mechanochromic behavior. As shown in Fig. 15 (bottom), the luminescence of the unground sample dramatically changes from green to yellow-orange after it has been completely ground. However, upon addition of a drop of acetonitrile the luminescence becomes bright again but is slightly different from the original luminescence. This difference can be assigned to the formation of a new crystalline phase of the material caused by the addition of the solvent. These interesting luminescence properties may originate from the one-dimensional stacking of the highly planar Pt^{II} complexes, where the bpy and **ImAm** ligands stack alternately.

Replacement of the NHP groups by NMe₂ substituents in the R₁ positions of the **ImAm** framework, as well as the addition of SiMe₃ in place of the H-atom at one of the R₂ positions, afford the ligand L²⁶. A slight modification of this ligand can also lead to L²⁷, in which 2,6-diisopropylphenyl (Dipp) is introduced on the amide N-atom instead of a Ph group. In the work by Zhou *et al.* these ligands were synthesized *in situ* during the formation of mononuclear and dinuclear complexes of Mg^{II}.⁵⁸ Although both ligand frameworks are rather similar, small changes in the synthetic conditions led to different

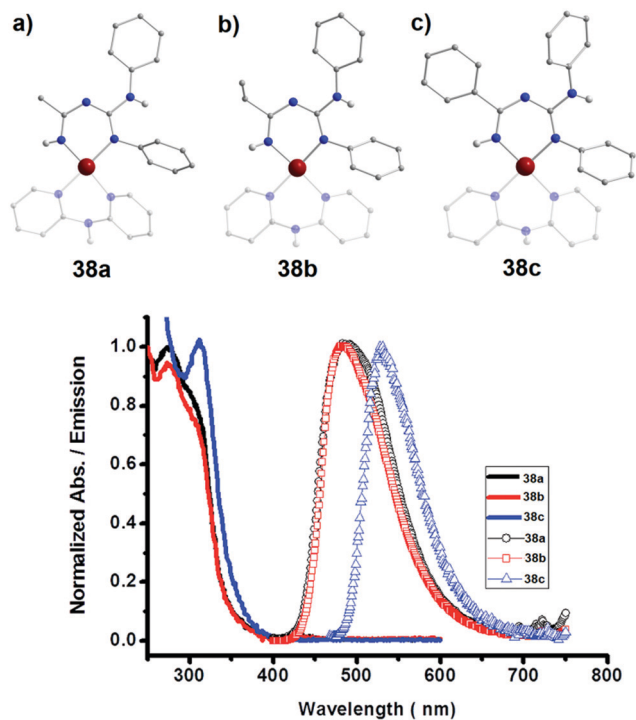


Fig. 14 Top: Molecular structures of **38a** (a), **38b** (b) and **38c** (c). For clarity reasons only the H-atoms of the nitrogen are shown, lattice solvent molecules and counter ions have been omitted, as well as partial transparency has been employed. Colour code: C: grey, N: blue, H: white, Pt^{II}: dark red. Reproduced using reported cifs. Bottom: Absorption in CH₂Cl₂ (solid lines) and the normalized solid-state emission spectra of **38a**, **38b** and **38c** at room temperature (icons with solid lines). Adapted with permission from ref. 56. Copyright 2014 Wiley.



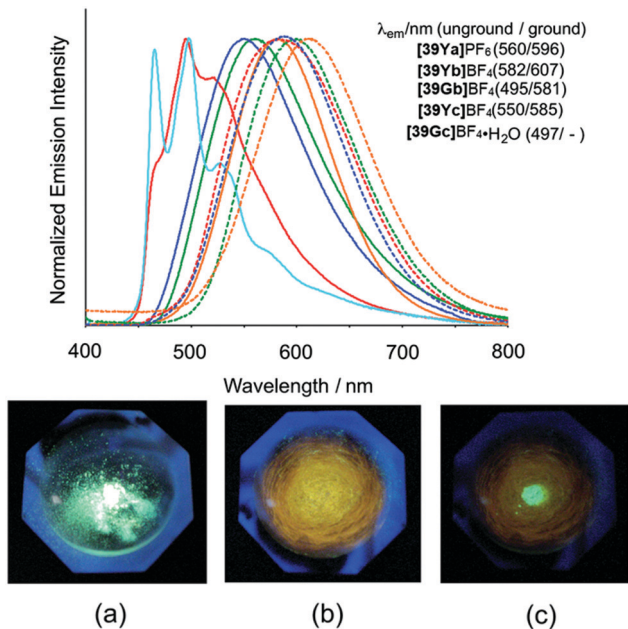


Fig. 15 Top: Normalized emission spectra of the Pt^{II} based complexes in the solid state at 295 K ($\lambda_{\text{ex}} = 355$ nm). Solid and dashed lines correspond to the spectra of unground and ground samples, respectively. Bottom: Photographic images of [39Gb]BF₄ in response to mechanical grinding under UV light irradiation (365 nm): (a) unground sample, (b) completely ground sample, (c) ground sample with a drop of acetonitrile. Reprinted with permission from ref. 57. Copyright 2014 Royal Society of Chemistry.

coordination motifs for each ligand, altering the nuclearity of the isolated complex. In the mononuclear complex [Mg^{II}(L²⁶)₂] (**40**), the ligand stabilizes the Mg^{II} center in a distorted tetrahedral geometry by acting as a chelate terminal ligand (U-shape fashion) (Fig. 16a). On the contrary, in the dinuclear [Mg^{II}Br₂(L²⁷)₂] (**41**), each Mg^{II} ion is bound to four N-atoms in a highly distorted square pyramidal geometry (a terminal bromo co-ligand completes the coordination sphere of each metal center) with both ligands adopting a *syn-syn* W-shape coordination motif, stabilizing the two Mg^{II} metal centers (Fig. 16b).

Apart from alkaline earth metals, the coordination chemistry of these ligands has also been explored with transition metals. Mononuclear complexes of Cu^{II} have been isolated with L²⁶: [Cu^{II}(L²⁶)₂] (**42**) and L²⁷: [Cu^{II}Cl(L²⁷)] (**43**) and [Cu^{II}(PPh₃)(L²⁷)] (**44**).^{58a,b} Both ligands coordinate on the metal center in a U-shape fashion, stabilizing the metal center in a square planar geometry in **42** and a trigonal planar geometry in **43** and **44**. The same group has also isolated the Zn^{II} analogue of **42**, which was demonstrated to possess catalytic properties for the ring-opening polymerization (ROP) of racemic-lactide.^{58c}

By replacing the NMe₂ groups with pyridine, L²⁸ can be formed. The obvious similarities to L²⁶ allows for a similar coordination behavior in the transition metal chemistry. Mononuclear complexes of this ligand with a generalized formula [M^{II}(L²⁸)₂] (where M = Zn^{II},^{58c} Fe^{II}⁵⁹ and Co^{II}⁵⁹) have been isolated. In all these cases, the ligands act as terminal chelates stabilizing the metal centers in a slightly distorted square planar

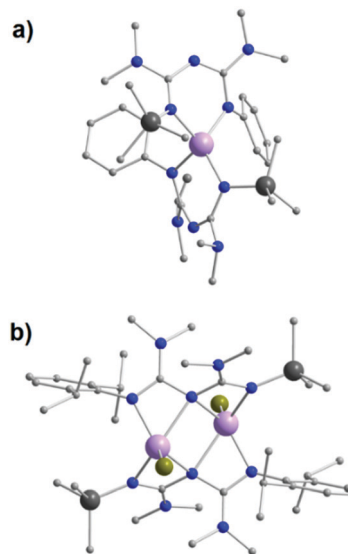
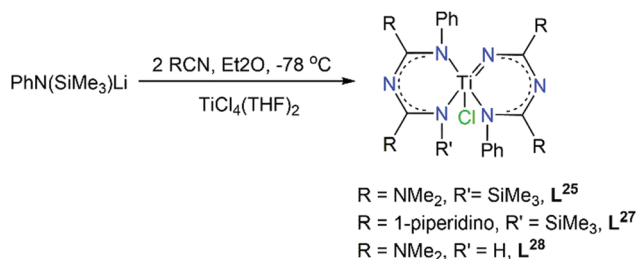


Fig. 16 The molecular structure of the complex **40** (a) and **41** (b).⁵⁸ For clarity reasons H-atoms have been omitted. Colour code: C: grey, N: blue, Br: dark yellow, Si: dark grey, Mg^{II}: mauve. Reproduced using reported cif's.

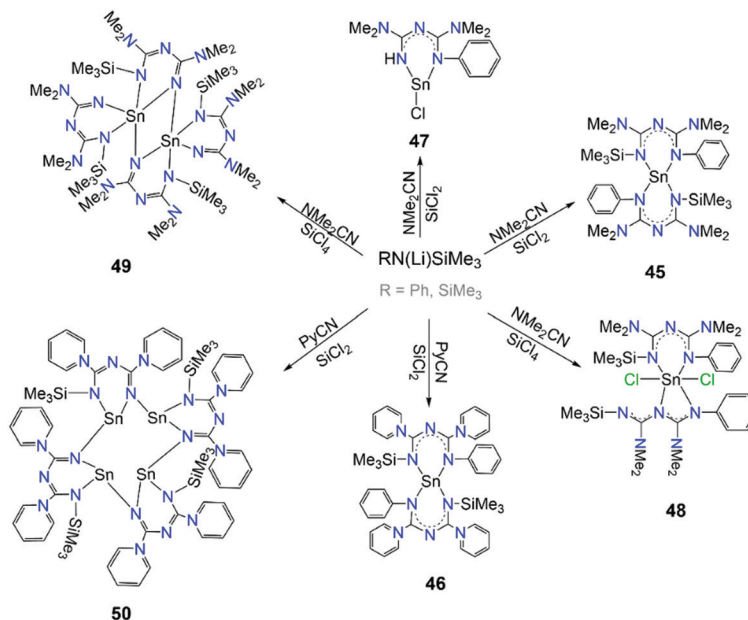
geometry, with the Zn^{II} analogue acting as a catalyst for the ROP of racemic-lactide.

The growing interest in asymmetric **ImAm** chelates and their ability to form stable complexes with transition metal ions, led Zhou and co-workers in studying the chemistry of L²⁶, L²⁷ and L²⁸ in Ti^{IV} complexes.⁶⁰ These complexes might also exhibit promising activities towards olefin polymerization. Starting from nitriles, the *in situ* synthesis of these ligands led to the isolation of a family of mononuclear Ti^{IV} complexes with the general formula [Ti^{IV}Cl(L^x)₂] (where x = 26, 28 or 29, respectively). In all three complexes the five-coordinate Ti^{IV} ion bonds to one chloro co-ligand and four nitrogen atoms from two **ImAm** ligands. Both ligands adopt the U-shape conformation, while the coordination geometry of the metal center can best be described as distorted trigonal bipyramidal. Surprisingly, one of the Ti–N bonds is significant shorter (average length 1.751(3) Å) than the other three distances (average length 2.075(6) Å), indicating that one ligand is in its dianionic form while the other is monoanionic (Scheme 18). Due to the potential activity towards olefin polymerization, the catalytic behavior of the Ti^{IV} complexes with L²⁸ and L²⁹ towards ethylene polymerization with methylaluminoxane (MAO) as co-catalyst was investigated,



Scheme 18 Synthesis of Ti^{IV} complexes with L²⁶, L²⁸ and L²⁹ as ligands.⁶⁰





Scheme 19 Synthetic routes for the preparation of complexes 45–49.⁶¹

demonstrating both complexes were active in ethylene polymerization.

The catalytic activity that the aforementioned complexes have shown inspired Guo *et al.* to explore the chemistry and catalytic activity of asymmetric **ImAm** ligands with Sn^{II}.⁶¹ For this purpose, several ligands such as L²⁶, L²⁸, L²⁹ and L³⁰ have been employed. These ligands were synthesized *in situ* from nitriles and by slight changes of the reaction conditions a family of mononuclear, dinuclear and tetranuclear complexes has been isolated. In Scheme 19, the synthesis of these complexes is given. In the mononuclear complexes [Sn^{II}(L²⁶)₂] (45) and [Sn^{II}(L²⁸)₂] (46), two monoanionic L²⁶ ligands are coordinated to the Sn^{II} metal center. Both L²⁶ are coordinated to the metal center in the U-shape coordination mode.^{61a} In the mononuclear [Sn^{II}Cl(L²⁹)] (47), the monoanionic L²⁹ coordinates in a similar fashion.^{61a} However, in 48 two different coordination motifs are observed, as one of the L²⁶ adopts a U-shaped conformation while the other a twisted W-shaped conformation. In this W-shaped mode, the N₃C₂ framework forms a twisted zigzag conformation.^{61b} Furthermore, higher nuclearities were obtained with L³⁰ and L³¹. In the dinuclear [Sn₂^{II}(L³⁰)₄] (49), dimerization is achieved through the unsubstituted N of the N₃C₂ framework of two dianionic L³⁰ moieties. The other two monoanionic L³⁰ act as terminal ligands to each Sn^{II} in a U-shape conformation.^{61b} By replacement of the NMe₂ groups with 1-piperidine, L³¹ leads to the formation of the tetramer [Sn₄^{II}(L³¹)₄] (50). Here, each dianionic L³¹ bridges two Sn^{II} ions through its unsubstituted N-atom.^{61a} Lastly, the catalytic activities of complexes 45, 47, 48 and 49 were tested. Complexes 45 and 47 were found to be efficient catalysts for the catalytic addition of phenylamine to *N,N'*-diisopropylcarbodiimide affording guanide,^{61a} while complexes 48 and 49 could be used as pre-catalysts in the addition of aniline or 2,6-diisopropylaniline to *N,N'*-diisopropylcarbodiimide.^{61b}

5. Conclusions, limitations and perspectives

From the results presented herein, it is clear that imidoamidines represent a versatile class of organic ligands for the construction of new mono- or polynuclear coordination complexes exhibiting a variety of coordination modes, coordination geometries and associated properties. The limited number of **ImAm** coordination complexes as well as their synthetic routes reported to date, reveals only the surface of this field has been scratched and deserves further exploration. Despite the Pinner synthetic procedure and other proposed synthetic routes by Muller, Häger, Hughes and their coworkers in the literature, the synthesis of **ImAm** chelates has not yet been fully explored and, therefore, holds great potential. In particular, these methods generally afford heavily *N*-substituted **ImAm** with bulky substituents at the terminal nitrogen atoms and/or strong electron-acceptor functionalities at the carbon atoms. Additionally, the overall yields reported are poor and the starting materials often are not commercially available, thereby adding more steps to their synthesis and limiting their further development. Another synthetic approach involves the use of metal mediated transformations, affording **ImAm**-based coordination complexes; however, this leaves little room for control during preparation of the **ImAm** framework. With that said, the recent report on an efficient, high yielding, one-pot synthesis of **ImAm**²¹ provides an avenue for creating a vast library of chelates with untapped potential. This therefore opens new perspectives in fields ranging from magnetic and luminescent materials to catalysis and biomedical applications.

Creative attention to both the organic and coordination chemistry of **ImAm** ligands can lead to the design of new **ImAm** frameworks and their metallo-assemblies. As an example, **P₂ImAm** affords two distinct coordination pockets; a tridentate



chelating group akin to terpyridine along with a bidentate chelating mode reminiscent of NacNac. These two coordination pockets afford stable five and six membered coordination rings, respectively, thus can be exploited to efficiently target coordination to various metal ions including mixed-metal systems. This unique ligand elegantly demonstrates the combination of two prolific ligand frameworks into a single chelate, which represents just one of the many **ImAm** to be recently explored. Overall, this review aims to highlight the versatility and vast potential of **ImAm**, a chelate like no other.

Conflicts of interest

The authors declare no conflict of interest.

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Notes and references

- (a) C. Pettinari, F. Marchetti and A. Drozdov, β -Diketones and Related Ligands, in *Comprehensive Coordination Chemistry*, Elsevier, Amsterdam, 2nd edn, 2004, pp. 97–115; (b) G. Aromi, P. Gamez and J. Reedijk, *Coord. Chem. Rev.*, 2008, **252**, 964–989; (c) A. J. Brock, J. K. Clegg, F. Li and L. F. Lindoy, *Coord. Chem. Rev.*, 2018, **375**, 106–133.
- (a) T. Kajiwara, M. Nakano, Y. Kaneko, S. Takaishi, T. Ito, M. Yamashita, A. Igashira-Kamiyama, H. Nojiri, Y. Ono and N. Kojima, *J. Am. Chem. Soc.*, 2005, **127**, 10150–10151; (b) A. M. Madalan, K. Bernot, F. Pointillart, M. Andruh and A. Caneschi, *Eur. J. Inorg. Chem.*, 2007, 5533–5540; (c) S. Wocadlo, W. Massa and J.-V. Folgado, *Inorg. Chim. Acta*, 1993, **207**, 199–206; (d) A. Kamiyama, T. Noguchi, T. Kajiwara and T. Ito, *Angew. Chem., Int. Ed.*, 2000, **39**, 3130–3132; (e) M. G. Cowan, R. G. Miller, P. D. Southon, J. R. Price, O. Yazaydin, J. R. Lane, C. J. Kepert and S. Brooker, *Inorg. Chem.*, 2014, **53**, 12076–12083; (f) S. Zebret, N. Dupont, G. Bernardinelli and J. Hamacek, *Chem. – Eur. J.*, 2009, **15**, 3355–3358; (g) S. Zebret, E. Torres, E. Terreno, L. Guénée, C. Senatore and J. Hamacek, *Dalton Trans.*, 2011, **40**, 4284–4290.
- (a) R. Bonnett, D. C. Bradley and K. J. Fisher, *Chem. Commun.*, 1968, 886–887; (b) R. Bonnett, D. C. Bradley, K. J. Fisher and I. F. Rendall, *J. Chem. Soc. A*, 1971, 1622–1627.
- J. E. Parks and R. H. Holm, *Inorg. Chem.*, 1968, **7**, 1408–1416.
- L. Bourget-Merle, M. F. Lappert and J. R. Severn, *Chem. Rev.*, 2002, **102**, 3031–3066.
- (a) Y.-C. Tsai, *Coord. Chem. Rev.*, 2012, **256**, 722–758; (b) S. P. Sarish, S. Nembenna, S. Nagendran and H. W. Roesky, *Acc. Chem. Res.*, 2011, **44**, 157–170; (c) M. Asay, C. Jones and M. Driess, *Chem. Rev.*, 2011, **111**, 354–396; (d) F. T. Edelman, *Coord. Chem. Rev.*, 2015, **284**, 124–205; (e) F. T. Edelman, *Coord. Chem. Rev.*, 2016, **318**, 29–130; (f) W. E. Piers and D. J. H. Emslie, *Coord. Chem. Rev.*, 2002, **233–234**, 131–155; (g) A. A. Mohamed, *Coord. Chem. Rev.*, 2010, **254**, 1918–1947; (h) V. T. Annibale and D. Song, *RSC Adv.*, 2013, **3**, 11432–11449; (i) D. Zhu and P. H. M. Budzelaar, *Dalton Trans.*, 2013, **42**, 11343–11354; (j) S. J. Malthus, S. A. Cameron and S. Brooker, *Coord. Chem. Rev.*, 2016, **316**, 125–161.
- (a) M. N. Kopylovich, A. J. L. Pombeiro, A. Fischer, L. Kloos and V. Y. Kukushkin, *Inorg. Chem.*, 2003, **42**, 7239–7248; (b) M. N. Kopylovich, M. Haukka, A. M. Kirillov, V. Y. Kukushkin and A. J. L. Pombeiro, *Chem. – Eur. J.*, 2007, **13**, 786–791; (c) N. A. Bokach, T. V. Kuznetsova, S. A. Simanova, M. Haukka, A. J. L. Pombeiro and V. Y. Kukushkin, *Inorg. Chem.*, 2005, **44**, 5152–5160; (d) P. V. Gushchin, M. R. Tyan, N. A. Bokach, M. D. Revenco, M. Haukka, M.-J. Wang, C.-H. Lai, P.-T. Chou and V. Y. Kukushkin, *Inorg. Chem.*, 2008, **47**, 11487–11500; (e) X. M. Guo, H. D. Guo, L. S. Fu, L. D. Carlos, R. A. S. Ferreira, L. N. Sun, R. P. Deng and H. J. Zhang, *J. Phys. Chem.*, 2009, **113**, 12538–12545.
- G. H. Sarova, N. A. Bokach, A. A. Fedorov, M. N. Berberan-Santos, V. Y. Kukushkin, M. Haukka, J. J. R. Fraústo-da-Silva and A. J. L. Pombeiro, *Dalton Trans.*, 2006, 3798–3805.
- A. A. Starikova, *Chem. Pap.*, 2018, **72**, 821–828.
- F. Kurzer, *Chem. Rev.*, 1956, **56**, 96–197.
- J. Baker, M. Kilner, M. M. Mahmoud and S. C. Wallwork, *J. Chem. Soc., Dalton Trans.*, 1989, 837–841.
- N. Heße, R. Fröhlich, I. Humelnicu and E.-U. Würthwein, *Eur. J. Inorg. Chem.*, 2005, 2189–2197.
- (a) N. Heße, R. Fröhlich, B. Wibbeling and E.-U. Würthwein, *Eur. J. Org. Chem.*, 2006, 3923–3937; (b) H. J. Breslin, M. J. Kukla, R. W. Tuman, M. C. Rebarchak and C. R. Bowden, *J. Med. Chem.*, 1993, **36**, 1597–1603; (c) I. Häger, R. Fröhlich and E.-U. Würthwein, *Eur. J. Inorg. Chem.*, 2009, 2415–2428; (d) N. C. Aust, A. Beckmann, R. Deters, R. Krämer, L. Terfloth, S. Warzeska and E.-U. Würthwein, *Eur. J. Inorg. Chem.*, 1999, 1189–1192; (e) H. Behrens, R. Fröhlich and E.-U. Würthwein, *Eur. J. Org. Chem.*, 2005, 3891–3899; (f) M. Buhmann, E.-U. Würthwein, M. H. Möller and U. Rodewald, *Angew. Chem., Int. Ed. Engl.*, 1994, **33**, 2337–2339; (g) A. R. Siedle, R. J. Webb, M. Brostrom, R. A. Newmark, F. E. Behr and V. G. Young, *Organometallics*, 2004, **23**, 2281–2286; (h) H. V. R. Dias and S. Singh, *Inorg. Chem.*, 2004, **43**, 7396–7402; (i) H. V. R. Dias, S. Singh and J. A. Flores, *Inorg. Chem.*, 2006, **45**, 8859–8861.
- M. N. Kopylovich and A. J. L. Pombeiro, *Coord. Chem. Rev.*, 2011, **255**, 339–355.
- D. A. Peak, *J. Chem. Soc.*, 1952, 215–226.
- H. Ley and F. Müller, *Ber. Dtsch. Chem. Ges.*, 1907, **40**, 2950–2958.
- A. W. Titherley and E. G. Hughes, *J. Chem. Soc.*, 1911, **99**, 1493–1510.
- H. C. Brown and P. D. Schuman, *J. Org. Chem.*, 1963, **28**, 1122–1127.



- 19 (a) F. C. Schaefer, I. Hechenbleikner, G. A. Peters and V. P. Wystrach, *J. Am. Chem. Soc.*, 1959, **81**, 1466–1470; (b) D. A. Safin, N. A. Tumanov, A. A. Leitch, J. L. Brusso, Y. Filinchuk and M. Murugesu, *CrystEngComm*, 2015, **17**, 2190–2195.
- 20 A. R. Siedle, R. J. Webb, F. E. Behr, R. A. Newmark, D. A. Weil, K. Erickson, R. Naujok, M. Brostrom, M. Mueller, S.-H. Chou and V. G. Young, *Inorg. Chem.*, 2003, **42**, 932–934.
- 21 (a) A. A. Leitch, I. Korobkov, A. Assoud and J. L. Brusso, *Chem. Commun.*, 2014, **50**, 4934–4936; (b) M. Yousaf, N. J. Yutronkie, R. Castañeda, J. A. Klein and J. Brusso, *New J. Chem.*, 2017, **41**, 12218–12224.
- 22 S. Hayashi, M. Furukawa, Y. Fujino and S. Yoshimatsu, *Chem. Pharm. Bull.*, 1969, **17**, 329–334.
- 23 (a) Z. Brzozowski, F. Saczewski and M. Gdaniec, *Eur. J. Med. Chem.*, 2000, **35**, 1053–1064; (b) L. Marin-Ocampo, L. A. Veloza, R. Abonia and J. C. Sepúlveda-Arias, *Eur. J. Med. Chem.*, 2019, **162**, 435–447; (c) S. Cascioferro, B. Parrino, V. Spano, A. Carbone, A. Montalbano, P. Barraja, P. Diana and G. Cirrincione, *Eur. J. Med. Chem.*, 2017, **142**, 328–375.
- 24 (a) M. Liu, L. Guo, S. Jin and B. Tan, *J. Mater. Chem. A*, 2019, **7**, 5153–5172; (b) C. Krishnaraj, H. S. Jena, K. Leus and P. Van Der Voort, *Green Chem.*, 2020, **22**, 1038–1071.
- 25 H. Wang, D. Jiang, D. Huang, G. Zeng, P. Xu, C. Lai, M. Chen, M. Cheng, C. Zhang and Z. Wanga, *J. Mater. Chem. A*, 2019, **7**, 22848–22870.
- 26 J. Artz, *ChemCatChem*, 2018, **10**, 1753–1771.
- 27 A. Loudet and K. Burgess, *Chem. Rev.*, 2007, **107**, 4891–4932.
- 28 F. Li, S. I. Yang, Y. Ciringh, J. Seth, C. H. Martin, D. L. Singh, D. Kim, R. R. Birge, D. F. Bocian, D. Holten and J. S. Lindsey, *J. Am. Chem. Soc.*, 1998, **120**, 10001–10017.
- 29 J. Bañuelos, *Chem. Rec.*, 2016, **16**, 335–348.
- 30 (a) J. M. Savard, R. Castañeda, B. Gabidullin and J. L. Brusso, *Polyhedron*, 2019, **162**, 8–13; (b) J. M. Savard, G. Brunet, A. Srinivasan, M. Murugesu and J. L. Brusso, *Dalton Trans.*, 2018, **47**, 14875–14879.
- 31 (a) C. P. Constantinides and P. A. Koutentis, *Adv. Heterocycl. Chem.*, 2016, **119**, 173–207; (b) P. J. Hayes, R. T. Oakley, A. Cordes and W. T. Pennington, *J. Am. Chem. Soc.*, 1985, **107**, 1346–1351.
- 32 P. J. Hayes, R. T. Oakley, A. W. Cordes and W. T. Pennington, *J. Am. Chem. Soc.*, 1985, **107**, 1346–1351.
- 33 R. T. Oakley, R. W. Reed, A. W. Cordes, S. L. Craig and J. B. Graham, *J. Am. Chem. Soc.*, 1987, **109**, 7745–7749.
- 34 R. T. Boéré, T. L. Roemmele and X. Yun, *Inorg. Chem.*, 2011, **50**, 5123–5136.
- 35 (a) A. A. Leitch, I. Korobkov, A. Assoud and J. L. Brusso, *Chem. Commun.*, 2014, **50**, 4934–4936; (b) E. Kleisath, N. J. Yutronkie, I. Korobkov, B. M. Gabidullin and J. L. Brusso, *New J. Chem.*, 2016, **40**, 4472–4479; (c) N. J. Yutronkie, P. Tami, S. Singh, E. Kleisath, B. M. Gabidullin, R. Davis and J. L. Brusso, *New J. Chem.*, 2017, **41**, 2268–2276; (d) N. J. Yutronkie, A. A. Leitch, I. Korobkov and J. L. Brusso, *Cryst. Growth Des.*, 2015, **15**, 2524–2532.
- 36 (a) K. L. M. Harriman, A. A. Leitch, S. A. Stoian, F. Habib, J. L. Kneebone, S. I. Gorelsky, I. Korobkov, S. Desgreniers, M. L. Neidig, S. Hill, M. Murugesu and J. L. Brusso, *Dalton Trans.*, 2015, **44**, 10516–10523; (b) N. J. Yutronkie, I. A. Kühne, I. Korobkov, J. L. Brusso and M. Murugesu, *Chem. Commun.*, 2016, **52**, 677–680; (c) K. L. Harriman, I. A. Kühne, A. A. Leitch, I. Korobkov, R. Clérac, M. Murugesu and J. L. Brusso, *Inorg. Chem.*, 2016, **55**, 5375–5383.
- 37 R. Norrestam, *Acta Crystallogr., Sect. C: Cryst. Struct. Commun.*, 1984, **40**, 955–957.
- 38 For example, see: (a) V. Duros, H. Sartzi, S. J. Teat, Y. Sanakis, O. Roubeau and S. P. Perlepes, *Inorg. Chem. Commun.*, 2014, **50**, 117–212; (b) M. M. Turnbull, M. Y. Wei and R. D. Willett, *J. Coord. Chem.*, 1995, **35**, 11–17; (c) M. N. Kopylovich, J. Lasri, M. F. C. Guedes da Silva and A. J. L. Pombeiro, *Dalton Trans.*, 2009, 3074–3084; (d) T. Kajiwara, A. Kamiyama and T. Ito, *Chem. Commun.*, 2002, 1256–1257; (e) M.-L. Tong, Y.-M. Wu, Y.-X. Tong, X.-M. Chen, H.-C. Chang and S. Kitagawa, *Eur. J. Inorg. Chem.*, 2003, 2385–2388; (f) D. R. Aris, J. Barker, P. R. Phillips, N. W. Alcock and M. G. H. Wallbridge, *J. Chem. Soc., Dalton Trans.*, 1997, 909–910; (g) C. R. G. Calona, A. S. Filatov, B. Pedro and E. V. Rybak-Akimova, *Organometallics*, 2019, **38**, 2512–2522; (h) W. J. Bland, R. D. W. Kemmitt, I. W. Nowell and D. R. Russell, *Chem. Commun.*, 1968, 1065–1066; (i) M. N. Kopylovich, A. M. Kirillov, E. A. Tronova, M. Haukka, V. Y. Kukulshkin and A. J. L. Pombeiro, *Eur. J. Inorg. Chem.*, 2010, 2425–2432; (j) J.-P. Zhang, Y.-Y. Lin, X.-C. Huang and X.-M. Chen, *J. Am. Chem. Soc.*, 2005, **127**, 5495–5506.
- 39 (a) Y.-Q. Xie, J.-J. Li, Y. Guo, Y.-M. Zhang and T.-B. Wei, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2013, **69**, m3; (b) G. Thiele, B. Wagner and S. Dehnen, *Eur. J. Inorg. Chem.*, 2015, 5329–5334.
- 40 I. Pernik, B. J. Maitland, A. Strasch and C. Jones, *Can. J. Chem.*, 2018, **96**, 513–521.
- 41 T. Bolaño, T. B. Gunnoe and M. Sabat, *Dalton Trans.*, 2013, **42**, 347–350.
- 42 (a) N. Q. Shixaliyev, A. M. Maharramov, A. V. Gurbanov, N. V. Gurbanova, V. G. Nenajdenko, V. M. Muzalevskiy, K. T. Mahmudov and M. N. Kopylovich, *J. Mol. Struct.*, 2013, **1041**, 213–218; (b) N. G. Shikhaliyev, A. M. Maharramov, V. M. Muzalevskiy, V. G. Nenajdenko and V. N. Khrustalev, *Acta Crystallogr., Sect. E: Struct. Rep. Online*, 2012, **68**, m1220–m1221.
- 43 N. Q. Shixaliyev, A. V. Gurbanov, A. M. Maharramov, K. T. Mahmudov, M. N. Kopylovich, L. M. D. R. S. Martins, V. M. Muzalevskiy, V. G. Nenajdenko and A. J. L. Pombeiro, *New J. Chem.*, 2014, **38**, 4807–4815.
- 44 M. N. Kopylovich, J. Lasri, M. F. C. Guedes da Silva and A. J. L. Pombeiro, *Eur. J. Inorg. Chem.*, 2011, 377–383.
- 45 (a) T. T. Ponduru, C. Qiu, J. X. Mao, A. Leghissa, J. Smuts, K. A. Schug and H. V. R. Dias, *New J. Chem.*, 2018, **42**, 19442–19449; (b) V. A. K. Adiraju, J. A. Flore, M. Yousufuddin and H. V. R. Dias, *Organometallics*, 2012, **31**, 7926–7932.
- 46 H. V. R. Dias, J. A. Flores, M. Pellei, B. Morresi, G. G. Lobbia, S. Singh, Y. Kobayashi, M. Yousufuddin and C. Santini, *Dalton Trans.*, 2011, **40**, 8569–8580.



- 47 K. Bakthavachalam and N. Dastagiri Reddy, *Organometallics*, 2013, **32**, 3174–3184.
- 48 (a) K. Bakthavachalam, B. Raghavendra and N. Dastagiri Reddy, *Appl. Organomet. Chem.*, 2017, **31**, e3833; (b) N. V. Kulkarni, A. Das, S. G. Ridlen, E. Maxfield, V. A. K. Adiraju, M. Yousufuddin and H. V. R. Dias, *Dalton Trans.*, 2016, **45**, 4896–4906.
- 49 (a) T. Kajiwarara and T. Ito, *Chem. Commun.*, 2002, 1256–1257; (b) M. N. Kopylovich, M. F. C. Guedes da Silva and A. J. L. Pombeiro, *Dalton Trans.*, 2009, 3074–3084; (c) V. Duros, H. Sartzi, S. J. Teat, Y. Sanakis, O. Roubeau and S. P. Perlepes, *Inorg. Chem. Commun.*, 2014, **50**, 117–121.
- 50 R. Castañeda, A. Hollingshead, B. Gabidullin and J. L. Brusso, *Cryst. Growth Des.*, 2017, **17**, 6572–6578.
- 51 R. Castañeda, K. L. M. Harriman, J. W. L. Wong, B. Gabidullin, M. Murugesu and J. L. Brusso, *Eur. J. Inorg. Chem.*, 2019, 963–972.
- 52 R. Castañeda, M. Rouzières, R. Clérac and J. L. Brusso, *J. Mater. Chem. C*, 2020, **8**, 4401–4407.
- 53 Z. Dubrawski, J. Heidebrecht, B. M. P. Lombardi, A. S. Hyla, J. Willkomm, C. L. Radford, J.-B. Lin, G. C. Welch, S. Ponnurangam, R. Roesler, D. E. Prokopchuk and W. E. Piers, *Sustainable Energy Fuels*, 2019, **3**, 1172–1181.
- 54 B.-B. Guo, Y.-J. Lin and G.-X. Jin, *Dalton Trans.*, 2017, **46**, 8190–8197.
- 55 (a) M. N. Kopylovich, Y. Y. Karabach, M. D. C. G. da Silva, P. J. Figiel, J. Lasri and A. J. L. Pombeiro, *Chem. – Eur. J.*, 2012, **18**, 899–914; (b) S. Jana, S. Khan, A. Bauzá, A. Frontera and S. Chattopadhyay, *J. Mol. Struct.*, 2017, **1127**, 355–360; (c) O. V. Nesterova, M. N. Kopylovich and D. S. Nesterov, *Inorganics*, 2019, **7**, 82; (d) E. V. Andrusenko, E. V. Kabin, A. S. Novikov, N. A. Bokach, G. L. Starova and V. Y. Kukushkin, *New J. Chem.*, 2017, **41**, 316–325.
- 56 I. I. Eliseev, P. V. Gushchin, Y.-A. Chen, P.-T. Chou, M. Haukka, G. L. Starova and V. Y. Kukushkin, *Eur. J. Inorg. Chem.*, 2014, 4101–4108.
- 57 T. Tanaka, R. Nouchi, Y. Nakao, Y. Arikawa and K. Umakoshi, *RSC Adv.*, 2014, **4**, 62186–62189.
- 58 (a) Q. Xie, H. Tong and M. Zhou, *Inorg. Chem. Commun.*, 2014, **44**, 37–40; (b) F. Li, L. Yan, H. Tong and M. Zhou, *Acta Crystallogr., Sect. E: Crystallogr. Commun.*, 2015, **71**, m54; (c) D. Tian, Q. Xie, L. Yan, H. Tong and M. Zhou, *Inorg. Chem. Commun.*, 2015, **58**, 35–38.
- 59 F. Li, X. Qiao, M. Wang, M. Zhou, H. Tong, D. Guo and D. Liu, *Polyhedron*, 2013, **52**, 639–644.
- 60 X. Wang, Z. Guo, H. Tong and M. Zhou, *J. Organomet. Chem.*, 2017, **828**, 10–15.
- 61 (a) Z. Guo, F. Liu, H. Tong, J. Chao, H. Wang and M. Zhou, *Polyhedron*, 2018, **151**, 273–278; (b) Z. Guo, D. Tian, Q. Yang, H. Tong, Y. Wang and M. Zhou, *Polyhedron*, 2019, **166**, 162–165.

