

# Materials Advances

Accepted Manuscript

This article can be cited before page numbers have been issued, to do this please use: S. Dutta, L. Negi and P. Munshi, *Mater. Adv.*, 2024, DOI: 10.1039/D4MA00444B.



This is an Accepted Manuscript, which has been through the Royal Society of Chemistry peer review process and has been accepted for publication.

Accepted Manuscripts are published online shortly after acceptance, before technical editing, formatting and proof reading. Using this free service, authors can make their results available to the community, in citable form, before we publish the edited article. We will replace this Accepted Manuscript with the edited and formatted Advance Article as soon as it is available.

You can find more information about Accepted Manuscripts in the [Information for Authors](#).

Please note that technical editing may introduce minor changes to the text and/or graphics, which may alter content. The journal's standard [Terms & Conditions](#) and the [Ethical guidelines](#) still apply. In no event shall the Royal Society of Chemistry be held responsible for any errors or omissions in this Accepted Manuscript or any consequences arising from the use of any information it contains.

# Multifunctional Single-component Organic Molecular Materials: Ferroelectricity, Negative Thermal Expansion, and Polymorphism

View Article Online  
DOI: 10.1039/D4MA00044BReceived 00th January 20xx,  
Accepted 00th January 20xx

DOI: 10.1039/x0xx00000x

Sanjay Dutta,<sup>a</sup> Lalita Negi<sup>a</sup> and Parthapratim Munshi<sup>\*,a</sup>

Organic materials with multifunctional properties, such as ferroelectricity, negative thermal expansion, and the phenomenon of polymorphism, have gained enormous interest from material scientists due to their unusual characteristics and unique applications. Moreover, research on single-component systems with such unusual properties is on the rise despite the necessity of satisfying strict criteria for designing such materials. This highlight provides an overview of the evolution of research on purely organic molecular ferroelectrics and thermo-responsive materials and the importance of studying polymorphism in such materials. Starting from their basics to the role of crystallographic symmetry in designing and characterizing, the mechanisms these special classes of materials follow and the possible use of various techniques for advancing these research areas are discussed here. This highlight is expected to advance materials science research and to motivate novice researchers to explore these key research areas.

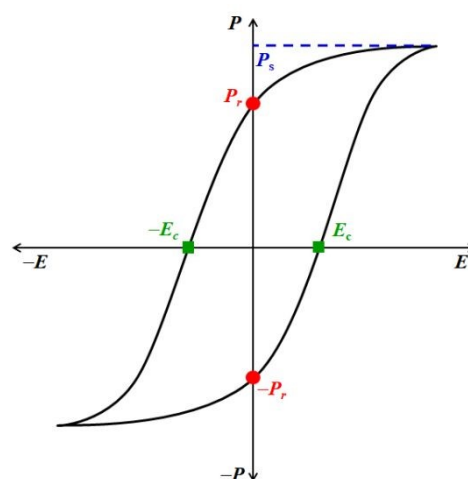
## 1. Introduction

Solid state science has made tremendous progress towards discovering multifunctional materials, among which ferroelectrics (FEs)<sup>1–3</sup> have been at the forefront. Simultaneously, thermo-responsive materials with unusual negative thermal expansion (NTE) properties have gained massive interest from material scientists due to their unique characteristics.<sup>4–7</sup> Moreover, polymorphism<sup>8</sup> has also drawn the attention of researchers as it plays a crucial role in governing material properties.

In this highlight, we aim to provide an overview of the research on purely organic molecular ferroelectrics and thermo-responsive materials and the role of polymorphism in such unusual materials. Initially, we introduce ferroelectricity and NTE properties, their applications, and the phenomenon of polymorphism. Then, we discuss the role of crystallography in designing and characterizing such materials and understanding their properties. Subsequently, we elaborate on the mechanisms these special classes of materials follow. After that, we focus on purely organic materials, especially the single-component systems. Before concluding, we discuss the possible use of various techniques for advancing these key research areas.

Ferroelectric is not related to the metal Fe (iron) as its name suggests, but it is analogous to ferromagnetism<sup>9</sup>. Like ferromagnetic systems, FE materials consist of multiple domains comprising electric dipoles. Domains<sup>10</sup> are small regions within the FE crystal where a group of electric dipoles are aligned in the same direction. Under the application of a strong external electric field ( $E$ ), these domains line up along its direction, resulting in net polarization ( $P$ ). This led to separating the positive and negative poles, similar to the magnetic dipoles in iron, where the north and south poles are created under a strong magnetic field. FEs are usually accompanied by temperature-triggered structural changes to undergo a reversible ferroelectric–paraelectric phase transition, resulting

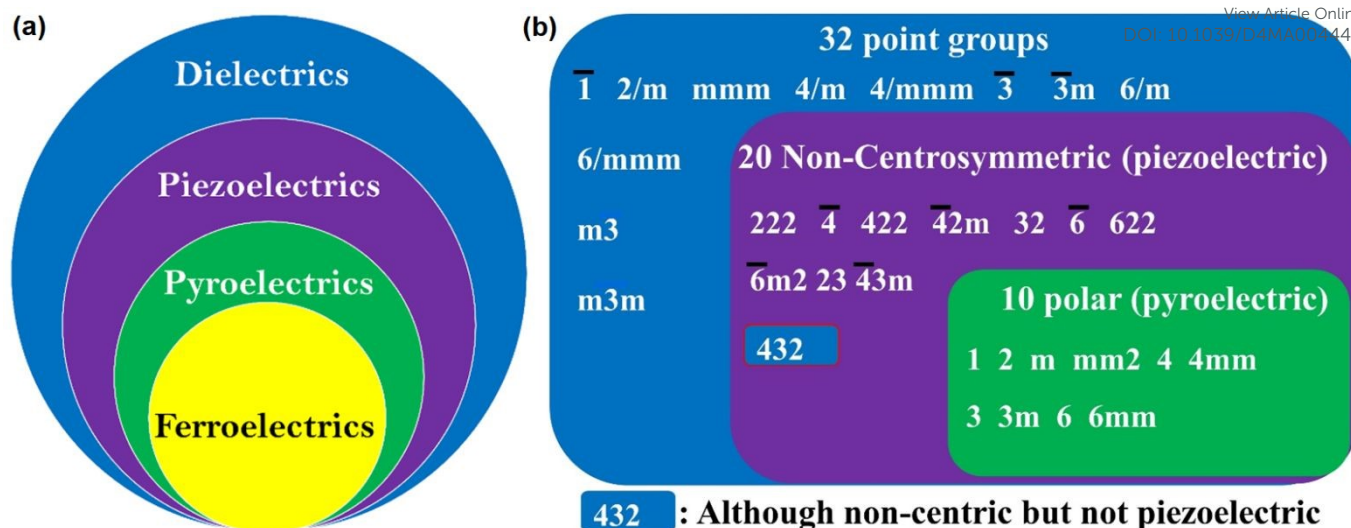
in the generation and vanishing of spontaneous polarization ( $P_s$ ). The net polarization in the FE materials is reversed or reoriented upon applying  $E$ . Moreover, due to their electrical bi-stability, ferroelectrics experience electrical displacement and display a hysteresis loop between the  $P$  and the  $E$  (**Figure 1**). These characteristics are paramount for fabricating storage devices such as FeFET<sup>2,3</sup> and FeRAM<sup>2,3</sup>. Ferroelectrics can also be used as capacitors because these materials obey Curie-Weiss law, and the dielectric constant rises rapidly during the ferroelectric to paraelectric transition, which is known as Curie temperature ( $T_c$ )<sup>1</sup>. The parameters that define the characteristics of FE materials are  $P_s$ , remanent polarization,  $P_r$ , and the coercive field,  $E_c$  (**Figure 1**).  $P_s$  value represents the maximum polarization induced in a FE material at a high  $E$ .  $P_r$  value indicates the amount of polarization the material can retain after removing  $E$ , i.e. at  $E = 0$   $\text{kV cm}^{-1}$ .  $E_c$  is the electric field required to reverse the polarization direction, i.e. to bring the net polarization of the material to zero ( $P = 0$   $\mu\text{C cm}^{-2}$ ). The figure of merit of FE materials is assessed based on their  $P_s$ ,  $P_r$ , and  $E_c$  values. Higher values of  $P_s$  and  $P_r$  and lower values of  $E_c$  signify superior characteristics of FE materials. Materials with low  $E_c$  find applications in low-voltage FE devices.<sup>11</sup>



**Figure 1:** Typical hysteresis loop ( $P$  vs  $E$ ) of a ferroelectric material and its associated characterization parameters.

<sup>a</sup>Multifunctional Molecular Materials Laboratory, Department of Chemistry, School of Natural Sciences, Shiv Nadar Institution of Eminence Deemed to be University, Delhi NCR, Uttar Pradesh – 201314, India  
E-mail: parthapratim.munshi@snu.edu.in





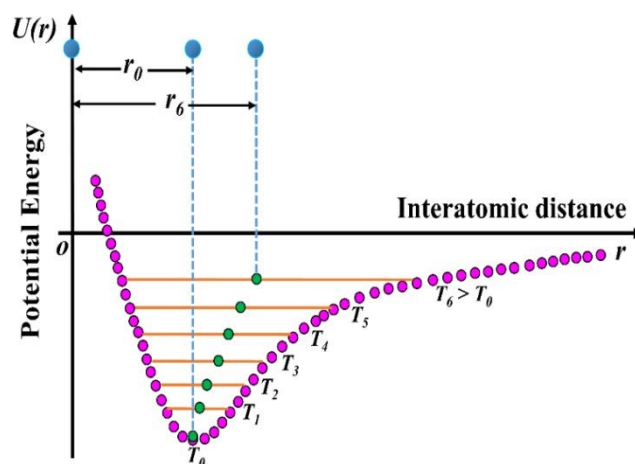
**Figure 2:** (a) The dielectric spectrum showing the super and sub-group relationship among the special class of electronic material of various physical properties, (b) Point group classification based on the symmetry.

As shown in **Figure 2a**, FEs are the lowest sub-group of dielectrics. They can be used in many applications due to their cross-correlation phenomena, such as piezoelectricity<sup>1</sup>, pyroelectricity<sup>1</sup> and second harmonic generation (SHG)<sup>12,13</sup>. In piezoelectric materials, due to electro-mechanical coupling, stress changes the electric  $P$ , and conversely,  $E$  produces strain<sup>1,14</sup>. Thus, these materials find applications in ultrasonic motion sensors, actuators, transducers, inkjet printers, and mechanical harvesters. The thermo-electric coupling leads to pyroelectric behaviour, resulting in the change in  $P$  due to uniform heating and cooling. This effect has been used in infrared detectors, thermal imaging and thermal energy harvesters<sup>1,15</sup>. Furthermore, the polar nature of FE materials also leads to electro-optic effects such as SHG, terahertz light emission and photovoltaic effects.

FEs had come to the foray of research as early as 1920 when *J. Valasek* discovered ferroelectricity in sodium potassium tartrate salt ( $\text{KNaC}_4\text{H}_4\text{O}_6$ ), also known as Rochelle salt<sup>16,17</sup>. However, for the next two decades, no FEs have been discovered. Consequently, the research community presumed it was the biggest serendipity of nature until the discovery of potassium dihydrogen phosphate- $\text{KH}_2\text{PO}_4$  (KDP) as a H-bonded FE in 1935<sup>18</sup>. Subsequently, the inorganic oxide barium titanate ( $\text{BaTiO}_3$ ) with perovskite structure was discovered as FE material in 1945<sup>19</sup>, which triggered the research in this important area. Thus, in the following years, the other perovskite structures<sup>20</sup> such as lead zirconate titanate ( $\text{Pb}(\text{Zr}_x\text{Ti}_{1-x})\text{O}_3$ ), bismuth ferrite ( $\text{BiFeO}_3$ ) and lithium niobate ( $\text{LiNbO}_3$ ) have come to the picture. These inorganic ferroelectrics have been a subject of active research for over 50 years and are used in many industrial and technological applications.

The usual phenomenon of expansion of solids with an increase in temperature is known as positive thermal expansion (PTE), which has been exploited in solid-state chemistry. The change in temperature of the material leads to different physical phenomena such as phase transition, heat-stimulated infrared detectors in pyroelectric materials<sup>15</sup>, and the generation of thermo-responsive switching materials<sup>21,22</sup>.

However, the expansion of solids upon heating hinders its applicability. For example, in large construction elements such as railway tracks, roads, bridges, and containers of liquefied natural gas and small precision devices such as fibre optic systems, organic light-emitting diode (OLED), organic field-effect transistors (OFET), thin-films, catalyst supports etc<sup>6,23,24</sup>. Nevertheless, the applications of multifunctional materials require not only robust physical phenomena but also controlled thermal expansion behaviour, as it enhances both lifetime and reliability. The effect of expansion upon heating can be correlated to the asymmetric shape of the Lennard-Jones potential energy well of a diatomic molecule (**Figure 3**). As the molecule is heated up, higher vibrational levels get populated, and the gap between the molecules increases<sup>6</sup>.



**Figure 3:** Schematic diagram of the potential energy well of paired atoms (adapted and modified from reference<sup>6</sup>).

The thermal expansion in solids is formally quantified with the coefficient of thermal expansion (CTE),  $\alpha$ , which can be calculated using the following equation:

$$\alpha_L = \frac{L_f - L_i}{L_i(T_f - T_i)} \quad \text{and} \quad \alpha_V = \frac{V_f - V_i}{V_i(T_f - T_i)}$$



Where  $\alpha_L$  and  $\alpha_V$  are linear and volumetric coefficients,  $L_f$  and  $V_f$  are the final length and final volume of the material of interest at a final temperature  $T_f$  and  $L_i$  and  $V_i$  are the initial length and initial volume at an initial temperature  $T_i$ , respectively. CTE is generally measured in parts per million per Kelvin ( $K^{-1}$ ), and its unit is  $K^{-1}$ . However, it is expressed in  $10^{-6} K^{-1}$  or  $MK^{-1}$  for conventional use. This equation holds well for the orthogonal crystal system in which the angles are perpendicular to each other.

Thermal expansion directions are identified by considering the *principal axes*, the set of orthogonal axes, X1, X2 and X3, along which thermal expansion varies linearly<sup>25</sup>. Thermal expansion is a directional property and depends on the symmetry of the crystal system. For cubic systems, expansion is isotropic in all three directions. Hence  $\alpha_V = 3\alpha_L$ . However, for the non-orthogonal system, the expansion is anisotropic and has different magnitudes of thermal expansion along the principal axes. The volumetric expansion,  $\alpha_V$ , can be calculated by adding  $\alpha_1$ ,  $\alpha_2$ , and  $\alpha_3$ , which are the CTE along each principal axis that are perpendicular to each other.

For inorganic compounds, the magnitude of  $\alpha_V$  lies in the range of 0-20  $MK^{-1}$ , and that of molecular materials usually lies in the range of 100-300  $MK^{-1}$ <sup>25</sup>. There are several techniques for monitoring thermal expansion in a solid. However, calculating CTE using unit-cell parameters obtained via single-crystal X-ray diffraction (SCXRD) is one of the most commonly used methods. Tools such as PASCAL<sup>26</sup> and Win\_Strain<sup>27</sup> are widely used to calculate CTE for non-orthogonal crystal systems.

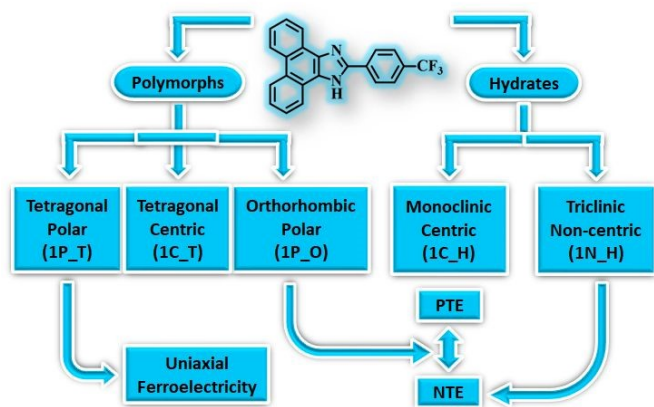
Contraction and expansion of materials upon heating and cooling, respectively, is termed NTE, a counter-intuitive phenomenon<sup>4-7</sup>. This special class of materials with such unusual properties can counter the PTE characteristics of the material<sup>28</sup>. However, the limited availability of NTE materials hinders the steady progress in this area of research. Thermal contraction was first documented in 1907 in quartz and silicates. Later, in 1948,  $NaZr_2(PO_4)_3$  (NZP)<sup>29</sup> was reported to have anisotropic NTE of  $\alpha_L = -0.4 MK^{-1}$  in the temperature range of 298–1273 K. The isotropic NTE effect was first discovered in 1996 by Sleight *et al.*<sup>30</sup> in  $ZrW_2O_8$ , which has a CTE value of  $-9 MK^{-1}$  over a wide temperature range of 0.3 to 1050 K. This has led to the significant progress of NTE-based studies in solid-state chemistry. Ice is another popular example of NTE material which displays NTE below 75 K due to low-energy transverse vibrational modes<sup>31</sup>. The emergence of the NTE phenomenon provides an excellent opportunity to control both physical properties and their thermo-responsive properties simultaneously, especially in the case of functional materials such as dielectric, piezoelectric, pyroelectric, magnetic and optical materials. Usually, NTE characteristics are noticed in inorganic systems. Most isotropic NTEs are found in open framework structures such as  $AM_2O_8$ <sup>32</sup>,  $AM_2O_7$ <sup>33</sup>,  $A_2M_3O_{12}$ <sup>34</sup>, zeolites<sup>35-37</sup> and ruthenium-based fluorides<sup>38-40</sup>. These materials experience NTE due to the transverse vibration of the bridging atom, which includes oxygen and fluorine. The other reports on the mechanism driving the NTE phenomenon are magnetic transition<sup>41</sup>, charge transfer<sup>42,43</sup>, and spontaneous volume ferro-electrostriction<sup>44,45</sup>, which are primarily prominent in inorganic functional materials.

NTE in FE materials is not uncommon<sup>6</sup>. However, the coexistence of FE and NTE has been reported predominantly in inorganic oxides<sup>46-48</sup>, benzo[18]crown-6 based metal complex<sup>49</sup>, and dabcoHCl.3H<sub>2</sub>O<sup>50</sup>. The report on the dabco complex demonstrated that the NTE occurs due to the rotation of water molecules, and the supramolecular architecture involving linear N-H...N bonds led to the FE behaviour. However, its anhydrous polar form divests these unusual properties. The NTE effect in FE materials generally involves multiple competing mechanisms, which are material-specific and depend on the complexity of the structural, electronic, and thermodynamic factors. NTE and FE properties are directional and are linked to temperature. While the anisotropy in a crystal structure is essential for the emergence of NTE and FE characteristics, the temperature variation can lead to a non-uniform response of the material due to the asymmetric arrangement of atoms in the crystal lattice. Moreover, under an external  $E$ , the macroscopic  $P$  reversal occurs in FE materials due to the reorientation of electric dipoles, which induce strains in the lattice with an increase or decrease of temperature, leading to lattice contractions or expansions. Furthermore, with the temperature change, the FE materials undergo structural phase transitions, which may trigger lattice softening in specific directions, leading to NTE along those directions. Further elucidation of the origin of the coexistence of these intriguing phenomena warrants more studies on versatile systems and delving into the role of microscopic interactions.

Polymorphism,<sup>8</sup> the phenomenon of the existence of a compound in multiple solid forms, has been widely studied in molecular materials for tuning their properties as it can significantly influence the material properties.<sup>51-55</sup> Polymorphs exhibit differences in structural, thermodynamic, and other physical properties. Often, polymorphs are categorised based on their molecular conformation, packing, colour, thermodynamic stability, etc. Accordingly, polymorphism is classified as conformational polymorphism, which arises due to the existence of different conformers of the same molecule; packing polymorphism, which arises due to the difference in molecular packing of the polymorphs in their crystal lattice; colour polymorphism, which arises due to the difference in colours of the polymorphic crystals, etc. Therefore, polymorphic materials can display distinct properties even with the slightest changes in their crystal structures<sup>56-58</sup>. Although polymorphs occur frequently in molecular crystals, their existence in organic materials with exotic properties is scarce<sup>54,59,60</sup>. Some of us have recently demonstrated that polymorphic forms exhibit distinct multifunctional properties. In the case of 2-(4-(trifluoromethyl)phenyl)-1H-phenanthro [9,10-d] imidazole, it exhibits FE and NTE properties in its various crystal forms – trimorphs and hydrates (Figure 4)<sup>61</sup>. One of the two polar forms exhibits FE characteristics, and the other displays a large reversible NTE below room temperature (RT). Further, the non-centrosymmetric hydrated form also experiences NTE below RT, but the centrosymmetric forms divest these properties. The other polymorphic materials with FE or NTE characteristics in organic materials have been highlighted in the following sections.







**Figure 4:** Multifunctional properties of polymorphic and hydrated forms of 2-(4-(trifluoromethyl)-phenyl)-1H-phenanthro [9,10-d] imidazole, a purely organic material.

## 2. Ferroelectricity

### 2.1 Crystallographic Perspective

In material science, crystallographic symmetry plays a critical role in identifying the structural phases of the solids and helps derive their structure-property relationships. The two basic symmetry operations in crystallography, rotation ( $n$ ) and reflection ( $m$ ), define the symmetry of a point in space. In crystallography, 32 point groups and 230 space groups defining the microscopic and macroscopic properties of the crystalline materials<sup>62</sup> are distributed among the 14 Bravais lattices and seven crystal systems, from higher to lower symmetry as cubic, hexagonal, rhombohedral (trigonal), tetragonal, orthorhombic, monoclinic and triclinic. Out of 32-point groups, 11 are centrosymmetric, and the rest 21 are non-centrosymmetric (Figure 2b). The 21 non-centrosymmetric point groups are further classified into 10 polar point groups and the rest into non-polar point groups. The lack of centre of inversion has a paramount effect on material properties, leading to numerous important physical properties in solid state science such as SHG, piezoelectricity, pyroelectricity and most importantly ferroelectricity<sup>63,64</sup>. FE materials must crystallize in one of the 10 polar point groups, and the corresponding paraelectric (PE) phase at a higher temperature must belong to one of the 32-point groups with a higher symmetry than the FE phase, as listed in Figure 2b. However, introducing the homochirality in a molecule can enhance the probability of a material crystallising in polar space ground. The chiral molecules crystallize in chiral point groups. There are 11 chiral point groups, out of which five are polar, so there is an increased probability (5/11 vs 10/32) that the molecule will crystallize in the polar point group<sup>65</sup>.

Based on the Curie symmetry principle<sup>66</sup>, the symmetry of the FE phase should be a maximal subgroup of the higher symmetric PE phase along which the direction of  $P_s$  exists. In light of this group-to-subgroup relationship, Aizu has identified 88 combinations of point groups for the PE to FE phase transition<sup>67</sup>. If the PE phase crystallizes in the centrosymmetric point group, the FE is considered reversible, where  $P_s$  can be reversed under  $E$ . If the PE phase exists in the polar

point group, then the FE is termed reorientable FE, where the direction of  $P_s$  is reoriented instead of reversing. Reorientable ferroelectrics are rare and fall into the category of relaxor-type FE. Moreover, under the guidance of the Curie symmetry-breaking principle, based on the point group symmetry of the known phase (PE), the symmetry and polarization direction of the other phase (FE) can be deduced using the Aizu relationship. Further, point group symmetry of the FE phase and its corresponding jump from low temperature and low symmetric phase to high temperature and high symmetry PE phase during symmetry breaking defines the number of allowed directions the  $P_s$  vector can have in FEs. These are then classified as (1) uniaxial (Rochelle salt, tri-glycine sulphate,  $\text{LiNbO}_3$ , etc.) and (2) multiaxial ( $\text{BaTiO}_3$ , PZT, etc.). In the former case, the  $P_s$  vector can reorient itself only in two polarization directions under  $E$ . For the latter case, the polarization vector can reorient itself in more than two polarization directions. As mentioned above, the point group of the FE phase is a subgroup of the PE phase. Point groups in the PE phase have various isomorphic subgroups whose polar direction corresponds to the direction of  $P_s$ , and the maximal one comprising the most symmetry elements along the polarization direction is the FE point group. Therefore, the number of polarization directions in the FE phase can be defined as  $n = N_{\text{PE}}/N_{\text{FE}}$ , where  $N_{\text{PE}}$  and  $N_{\text{FE}}$  represent the sum of symmetry elements in the PE and FE point groups, respectively. Thus, when  $n = 2$ , the FEs are uniaxial with only two opposite polarization directions, and more than two polarization directions characterize multiaxial FEs. The greater the symmetry change between the PE and FE phase, the higher the equivalent polarization directions a FE material will have. For example, in the case of  $\text{LiNbO}_3$ , both the PE space group,  $R\bar{3}c$  ( $N_{\text{PE}}=4$ ), and the FE space group,  $R3c$  ( $N_{\text{FE}}=2$ ), belong to the trigonal crystal system. Closer lattice symmetries led to arousal of uniaxial nature.<sup>68</sup> Similarly, diisopropylammonium bromide, molecular FE with a high  $T_c$  (426 K) and a large  $P_s$  ( $23 \mu\text{C cm}^{-2}$ ), is also uniaxial with an Aizu notation of  $2/mF2$ .<sup>69</sup> FE transitions from a cubic crystal system to one with much lower symmetry are generally multiaxial, and the  $n$  value can be as high as 48. For instance, in  $\text{BaTiO}_3$ , which is a multiaxial system, the successive phase transitions from cubic  $Pm\bar{3}m$  ( $N_{\text{PE}} = 48$ ) to tetragonal  $P4mm$  ( $N_{\text{FE}} = 8$ ) at 393 K, to orthorhombic  $Amm2$  ( $N_{\text{FE}} = 4$ ) at 278 K, and then to  $R3m$  ( $N_{\text{FE}} = 6$ ) at 183 K, led to multiaxial ferroelectric behaviour with 6, 12, and 8 equivalent polarization directions.

### 2.2. Ferroelectricity Mechanisms

Based on the type of chemical bonding, the ferroelectrics are classified as H-bonded ferroelectrics, e.g.  $\text{KH}_2\text{PO}_4$ , Rochelle salt, tri-glycine sulphate, etc., and inorganic oxides having a perovskite structure, e.g.  $\text{BaTiO}_3$ , PZT, etc. The third classification is based on the (1) order-disorder, (2) displacive, and (3) proton transfer mechanisms.

#### 2.2.1 Order-disorder

In this type, a thermally disordered state reorients to an ordered state. In Figure 5a, this mechanism is represented in the case of sodium nitrite ( $\text{NaNO}_2$ )<sup>1</sup>, where a permanent dipole of the nitrite ion



(NO<sub>2</sub><sup>-</sup>) generates  $P_s$ , and reorientation of dipoles upon application of an external  $E$  originates the ferroelectricity. In the FE phase, the orientation of the NO<sub>2</sub><sup>-</sup> is ordered, and it becomes disordered at a higher temperature, leading to the cancellation of the dipole moment and resulting in the corresponding PE phase (Figure 5a). Amine-based organic salts, such as diisopropylammonium bromine (C<sub>6</sub>H<sub>16</sub>NBr)<sup>69</sup> and chloride (C<sub>6</sub>H<sub>16</sub>NCl)<sup>70</sup>, which are low molecular weight FEs, belong to the order-disorder type. The origin of ferroelectricity in these systems is due to the ordered arrangement of the diisopropyl ammonium cation (C<sub>6</sub>H<sub>16</sub>N) in the FE phase upon cooling down from its PE phase, where the cation exhibits rotational disorder. A similar scenario was observed in the case of DABCO-based organic salts<sup>71</sup>.

### 2.2.2 Displacive

If the origin of ferroelectricity is due to the relative displacement of ions or molecules, the ferroelectrics are categorized as displacive type. The classic example is BaTiO<sub>3</sub><sup>19</sup>, having an ABX<sub>3</sub> type perovskite structure (Ba<sup>2+</sup> as A, Ti<sup>4+</sup> as B and O<sup>2-</sup> as X), and its cubic structure at high temperatures undergoes three successive FE phase transitions. Each phase involves slight distortion from the high symmetry phase (Figure 5b). The net displacement of the cations with respect to the oxygen octahedra results in  $P_s$  in the FE phases. Displacive-type FEs have also been discovered in metal-organic frameworks (MOF). The (15-crown-5)-bis (nitrate-O, O')-calcium (II)<sup>72</sup> is an excellent example of low-temperature molecular FE where distortion of the coordination environment of Ca<sup>2+</sup> induces ferroelectricity.

### 2.2.3 Proton transfer

In this type, the labile protons on hydrogen bonds activate the FE ordering within the system. These are also called KDP type because the collective proton shuttling in the O-H...O bonds leads to the origin of ferroelectricity (Figure 5c). Both order-disorder and displacive

mechanisms are not mutually exclusive as they can exist together, and the KDP is a prime example where proton ordering triggers the ferroelectricity, but main contribution to  $P_s$  comes from the displacement of phosphate ion (PO<sub>4</sub><sup>3-</sup>) and potassium (K<sup>+</sup>) ion<sup>18,73</sup>. A similar mechanism is proposed in the case of DABCO-based organic salt<sup>71</sup> of [H<sub>2</sub>dabco]·[2CB] ([H<sub>2</sub>dabco]<sup>2+</sup>=diprotonated 1,4-diazabicyclo [2.2.2] octane, 2CB<sup>-</sup>=2-chlorobenzoate), which undergoes FE to PE transition at 323 K, leading to both displacement of proton position and simultaneously disordering of the [H<sub>2</sub>dabco]<sup>2+</sup> cation. In the PE phase, the [H<sub>2</sub>dabco]<sup>2+</sup> cation is rotationally disordered and lies at the symmetric centre of the trimer. It freezes in an ordered state upon cooling and deviates toward a 2CB<sup>-</sup> anion at one end along the H-bond. This collective displacement and ordering of the cations lead to the origin of polarization in the FE phase.

Further, an optically controlled polarization switching mechanism has been explored in the domain of inorganic FEs. Various photo-excited electronic effects such as flexoelectric and photovoltaic have been studied to gain photocontrol in such material. Light-driven effects, such as the flexoelectric effect due to photoexcitation of electronic effects to control optical polarization, have been seen in the case of SrTiO<sub>3</sub><sup>74</sup>. Some specific organic molecules have recently been exploited to generate polarization using light as a stimulus<sup>75</sup>.

### 2.3. Organic Ferroelectrics

Though the first FE material, Rochelle salt<sup>16,17</sup> is an organic salt, the past century has seen the dominance of inorganic ceramics, and the number of FEs that have come to the fore are mainly inorganic oxides and polymeric systems. Traditionally, inorganic ceramics have been used for many practical electronic applications due to their superior FE characteristics and robust piezoelectric behaviour. However, using inorganic compounds containing heavy elements has

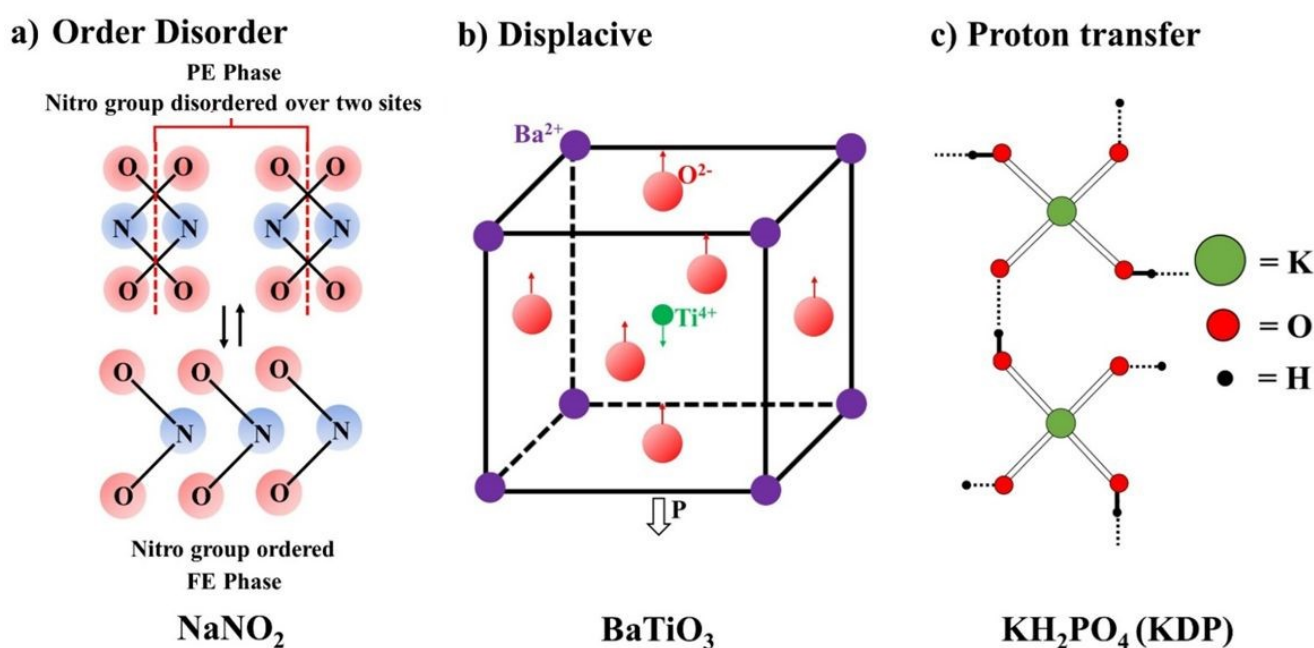


Figure 5: Different types of mechanisms that lead to the origin of ferroelectricity polarization.



become a serious global concern, especially regarding health and the environment. Thus, the scientific community was forced to look for alternatives, leading to the resurgence of organic FEs in the past decades due to their benign nature, mechanical flexibility, lightweight, environment-friendliness, cost-effective production and easy processing<sup>76–81</sup>. Further, organic electronics provide the design flexibility to tune the properties and enhance their performance and durability. Accordingly, organic FE devices can be portable, flexible, wearable, biocompatible and usable as economical medical devices.

The order-disorder mechanism in multicomponent systems led to a significant development of molecular FEs. Sachio *et al.* developed several multicomponent FEs based on acid-base supramolecular co-crystals with heteronuclear hydrogen bonds, which undergo the abovementioned mechanism. Some examples are phenazine-chloranilic acid<sup>82</sup>, and binary combinations of 2,5-dihalo-3,6-dihydroxy-p-benzoquinones with bases such as phenazine<sup>83</sup>, 5,5-dimethyl bipyridine<sup>84</sup>, 2,3-di(2-pyridinyl) pyrazine<sup>85</sup>. The advent of ferroelectricity in these systems amounts to the degree of proton transfer, which strongly depends on the acidity and basicity of the cofomers. Based on the combination, there could be two types of supramolecular systems, neutral and ionic. Harada *et al.* used globular-shaped molecules to design molecular FEs, especially in plastic crystals<sup>86,87</sup>. The globular molecule structure in plastic crystals can easily undergo order-disorder transition due to the low rotational barrier, which led to the development of some multiaxial FEs such as tetramethylammonium tetrachloroferrate(III)<sup>87</sup> and tetramethylammonium bromotrichloro ferrate(III)<sup>87</sup>.

Further, Xiong *et al.* developed metal-free perovskite-based order-disorder type FE<sup>88</sup>. Typically, perovskites have a three-dimensional (3D) ABX<sub>3</sub> structure in which the A cations are located in the 3D corner-sharing framework of BX<sub>6</sub> octahedrons (B represents the other cation, and X is an anion). Such perovskite represents one of the most pursued forms of isotropic ferroelectrics. The globular shape molecules based on DABCO derivatives have been used to design such molecular FEs. MDABCO-NH<sub>4</sub>I<sub>3</sub> is one such derivative, which displays FEs with multiaxial characteristics above RT. In this case, the large organic cations (MDABCO), small NH<sub>4</sub><sup>+</sup>, and halide anions (I<sup>-</sup>) occupy the A, B, and X-sites, respectively. The B-site cations are also involved in the NH...X hydrogen bonds with the surrounding X-site anions. The phase transition involves a cubic PE structure and a rhombohedral FE structure at 448 K, accompanied by a large off-centred displacement of the A- and B-site ions. This led to a multiaxial nature with Aizu notation<sup>67</sup> of *432F3*, having eight polarization axes along [111]. The high  $P_s$  of 22  $\mu\text{C cm}^{-2}$  is observed due to the synergistic effect of two distinct mechanisms, ionic displacement and order-disorder orientation. The other examples are diisopropylammonium bromide<sup>69</sup>, diisopropylammonium chloride<sup>70</sup> and guanidinium perchlorate<sup>89</sup> with  $P_s$  of 23, 8.9 and 8.4  $\mu\text{C cm}^{-2}$ , respectively.

Furthermore, efforts have been made to discover proton transfer mechanism-based organic FEs via database mining.<sup>90</sup> The study also proposed density functional theory-based prediction of  $E_c$  in such systems.

Finally, the development of isotropic molecular FEs eased the fabrication of thin film devices. Conventional organic FEs with highly anisotropic crystal structures cannot work unless their crystal orientation leads to a uniaxial polarization in the desired direction. In contrast, the nearly isotropic characters of possible polarization vectors guarantee efficient switching, even in randomly oriented crystals such as polycrystalline thin films and compacted pellets.

## 2.4. Single-component organic ferroelectrics

The multicomponent systems comprised of salts<sup>66,71,92</sup> and co-crystals<sup>81,93</sup> have certain disadvantages, which can be overcome through single-component organic ferroelectrics (SCOFs). The concern regarding salt systems is that they often contain metal ions, which leads to toxicity. However, the addition of metal ions may enhance  $P_s$  and  $T_c$ . Co-crystals, although benign, suffer from lower thermal stability and shelf life, hindering their applicability. These challenges have been overcome in the case of SCOFs. However, in SCOFs, due to the thermal robustness of the FE phase, the  $T_c$  is usually hidden beyond its thermal stability. Various types of SCOFs reported in the literature are summarized below (Figure 6).

### 2.4.1 PTM-type ferroelectrics

Tautomeric systems, such as keto-enol and enamine-imine, which incorporate D and A in the same molecule, can be used to design SCOFs (Figure 7a). This idea was first proposed by Haddon *et al.* in 1982 when H-bonded tautomeric systems could be used as model systems for memory application devices<sup>94</sup>. The first prototype study for such systems was squaric acid<sup>95–97</sup> containing the  $\beta$ -diketone enol (enolone, HO–C=C–O). Squaric acid undergoes tautomerization between the keto and enol forms (Figure 7a), exhibiting strong coupling between the proton and  $\pi$ -electron system. This type of prototropic system reinforces intra or intermolecular hydrogen bonds synergistically with the augmentation of the  $\pi$ -system delocalization, defined as resonance-assisted H-bond (RAHB)<sup>98–100</sup>. The FE systems developed under this principle are defined as proton tautomerization type ferroelectrics, and the corresponding mechanism is known as proton tautomerism mechanism (PTM)<sup>80</sup> as illustrated in Figure 7b.

In addition to proton tautomerization, the other prime factors for designing FE systems are the presence of inherent molecular symmetry higher than C<sub>1</sub><sup>101</sup> and the ability of a polar structure to have hidden pseudosymmetry, which persists in the PE phase. The concept of pseudosymmetry in single-component systems is based on the intrinsic mechanism of ferroelectricity, where there is a subtle deviation of polar structure in the FE state from the centrosymmetric structure, thus ensuring easier polarization switching. This principle has been utilised in the past by Zikmund *et al.*<sup>102</sup> to predict the potential FE candidate. They discovered weak ferroelectricity in cyclohexane-1,1 $\alpha$ -diacetic acid<sup>103</sup>. The hidden pseudosymmetry in the proton transfer type FEs can be identified by the global crystal symmetry of the proton whose position defines the crystal polarity. Local pseudosymmetries in the crystal structures can be identified using the ADDSYM option in the PLATON program<sup>95</sup>. Thus, based on





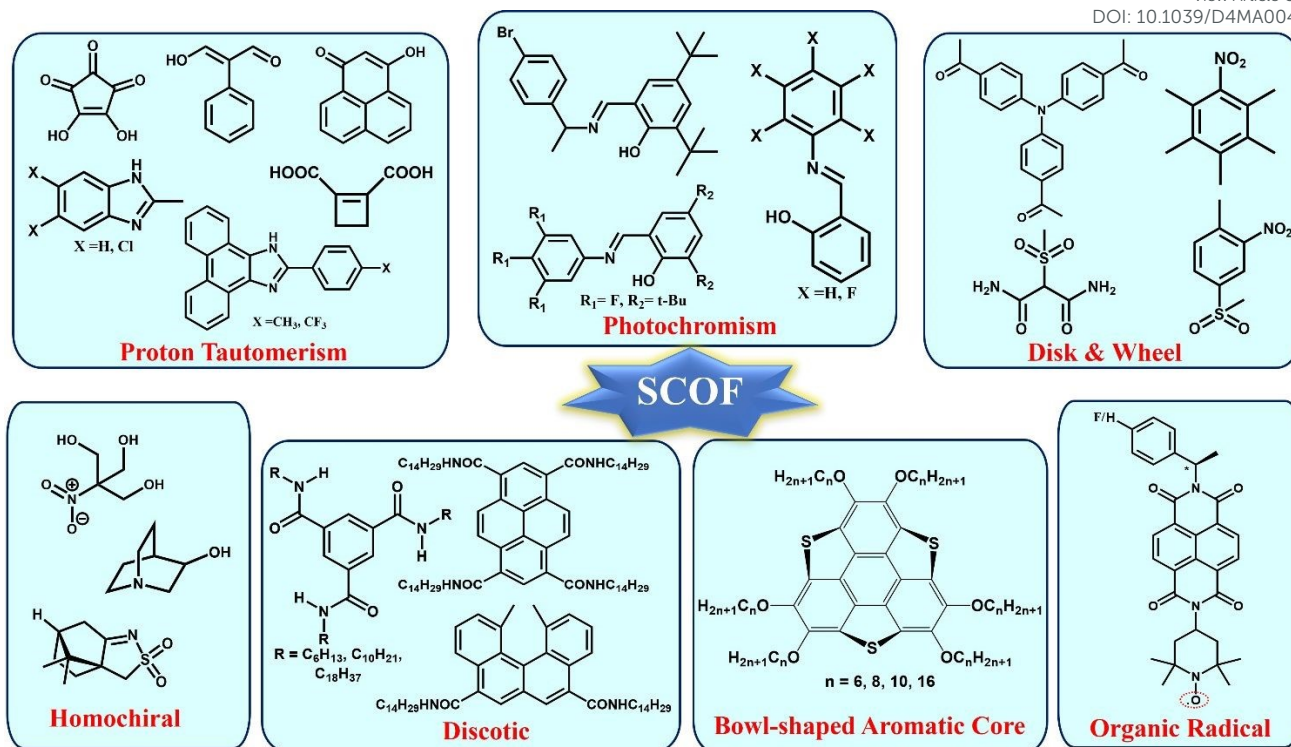


Figure 6: Examples of different types of single-component organic ferroelectrics.

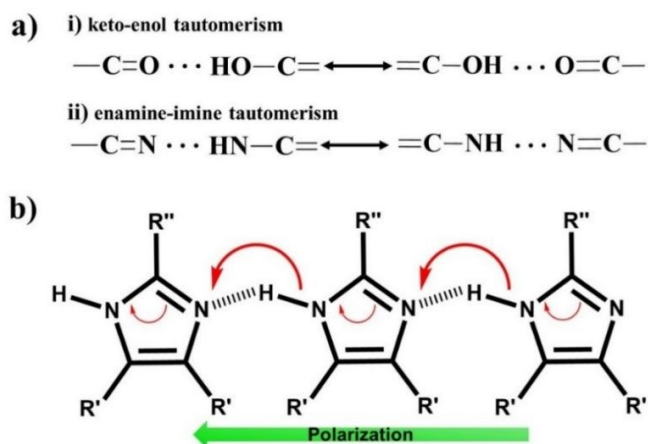


Figure 7: (a) Tautomeric system and (b) proton tautomerism mechanism in enamine-imine based single-component organic ferroelectrics.

this initial hypothesis, researchers started to explore systems consisting of tautomeric units to characterize them as FEs. The breakthrough study in this direction was achieved in 2010 when Sachio *et al.* discovered croconic acid (CA) as a FE material<sup>77</sup>. CA is a black organic dye<sup>104</sup>, a 5-member ring of oxocarbon acids ( $\text{H}_2\text{C}_n\text{O}_n$ ) that is strongly acidic because its conjugate base ( $\text{C}_5\text{O}_5^{2-}$ ) dianions formation is favoured due to delocalization of the  $\pi$ -electron within the ring. This amphoteric nature of CA led to strong RAHB<sup>98–100</sup> formation, thus undergoing a keto-enol type of proton tautomerism and contributing to a large molecular dipole moment, which sums up to give rise to the net polarization of  $30 \mu\text{C cm}^{-2}$  in its crystal form. In

addition to the enhanced polarization, low  $E_c$  and robust thermal stability ( $\sim 450\text{K}$ ) observed in CA<sup>77</sup> can be correlated to the strong coupling between the proton and the  $\pi$ -electron cloud. The concept of hidden pseudosymmetry in the case of CA is illustrated in Figure 8. The pseudo-mirror symmetry, which orthogonally crosses the pentagon of the CA ring, leads to the formation of the hypothetical PE phase (proton in disorder state) with space group  $Pcam$ , one of the minimal supergroups of the polar space group  $Pca2_1$  of the FE phase. The other systems comprising  $\beta$ -diketone enol  $\text{O=C—C=C—OH}$  moieties as a simpler prototype for the FE tautomerism are 2-phenylmalondialdehyde and 3-hydroxyphenalenone<sup>78</sup>.

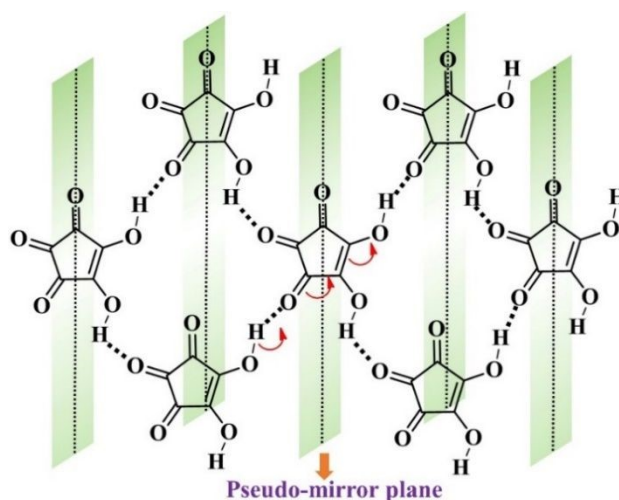
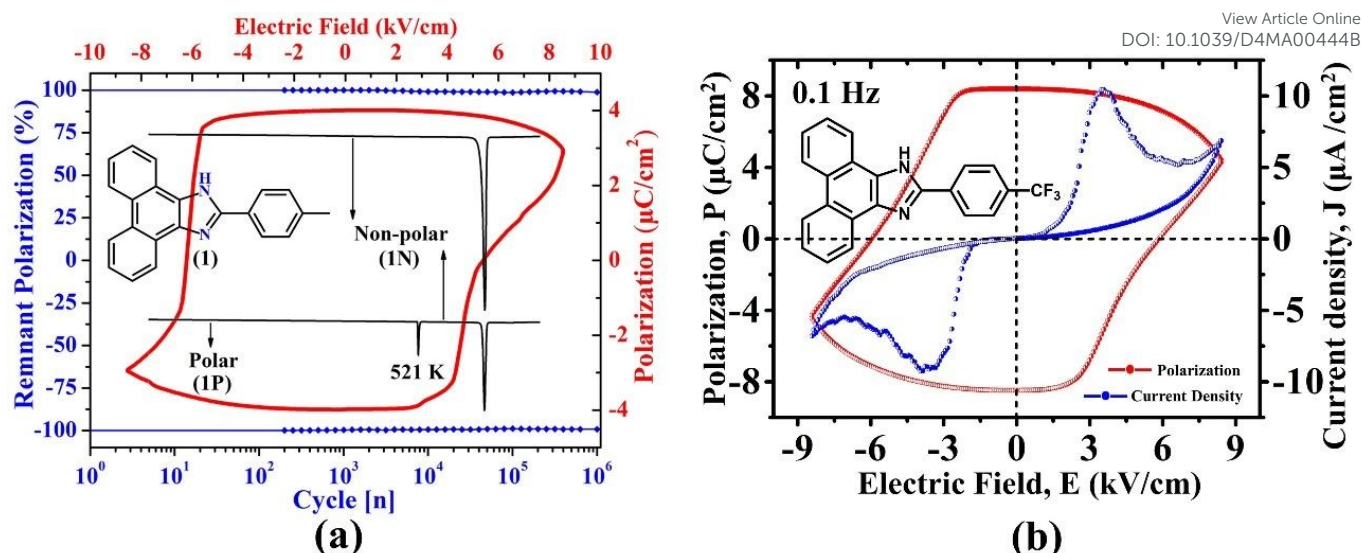


Figure 8: Pseudo-mirror symmetry in croconic acid with resonance-assisted hydrogen bonds.







**Figure 9:** (a) Depiction of ferroelectricity, thermal stability and phase transition behaviour in 2-(p-tolyl)-1H-phenanthro [9,10-d] imidazole.<sup>54</sup> (b) Ferroelectric loop ( $P$  vs  $E$ ) and corresponding fatigue measurements in 2-(4-(trifluoromethyl) phenyl)-1H-phenanthro[9,10-d] imidazole.<sup>61</sup>

Ferroelectricity was also discovered in imidazole-based derivatives comprising the enamine unit (Figure 7b). These enamine-imine motifs in imidazole exhibit amphoteric behaviour, thus forming a proton D-A type molecular assembly based on linear N-H...N hydrogen bonds. Imidazole-based SCOFs were reported in the cases of two benzimidazole derivatives, namely, 2-methylbenzimidazole<sup>79</sup> and 5,6-dichloro-2-methylbenzimidazole<sup>79</sup>. Further, Fraser *et al.* discovered ferroelectricity in halogen-substituted imidazoles, where they confirmed the ferroelectricity based on local polarization switching using the piezo force microscopic (PFM) technique<sup>105</sup>.

Recently, some of us reported SCOF in an enamine-based motif, 2-(p-tolyl)-1H-phenanthro [9,10-d] imidazole, which exists in three polymorphic forms<sup>54</sup>. The polar crystal form of this environmentally benign material not only displays notable  $P_s$  of  $3.0 \mu\text{C cm}^{-2}$ , highly stable (98%)  $P_r$  of  $3.9 \mu\text{C cm}^{-2}$ , and low  $E_c$  ( $6.2 \text{ kV cm}^{-1}$ ), it perseveres FE phase up to a record-high temperature of  $\sim 521 \text{ K}$  (Figure 9a). Subsequently, the crystal undergoes an irreversible phase transition to form non-polar (1N,  $P4_2/c$ ) and centrosymmetric ( $P2_1/n$ ) polymorphs having distinct molecular packing in their crystal lattices. These polymorphic transformations suggested that the material loses its FE characteristics beyond  $\sim 521 \text{ K}$  and remains non-ferroelectric. Nevertheless, this was the thermally most stable purely organic FE crystal then. Moreover, the  $E_c$  achieved in this case is much lower ( $6.2 \text{ kV cm}^{-1}$ ) than those reported for the well-known CA ( $14 \text{ kV cm}^{-1}$ ) and chloro derivative of benzimidazole ( $60 \text{ kV cm}^{-1}$ ) at  $373 \text{ K}$ . Our study demonstrated that the migration from benzimidazoles to phenanthroimidazoles with extended benzene rings helps maintain the FE phase at a much higher temperature. Further, the study highlighted that although the PTM triggers the generation of polarization and the polarity switching via RAHB, the low value of  $E_c$  is achieved due to the relative displacements between the two crystallographically independent molecules in its crystal lattice. This study was the first to report displacive type FE in SCOF. Otherwise, it usually occurs in inorganic oxides.

Further, the  $P_s$  was computationally calculated using Berry phase formalism and the DFT method with the generalized gradient approximation in the Perdew–Burke–Ernzerhof (PBE) form as implemented in the Vienna ab initio simulation package (VASP). The estimated  $P_s$  of  $\sim 3.8 \mu\text{C cm}^{-2}$  compares well to that obtained from the P-E loop measurement at RT.

Furthermore, some of us discovered that the tetragonal polymorphic form ( $1P_T$ ,  $P4_1$ ) of (2-(4-(trifluoromethyl)phenyl)-1H-phenanthro[9,10-d]imidazole)<sup>61</sup>, a triflate-substituted phenanthroimidazole, is a SCOF (Figure 9b). In this case, the in-crystal polarization reversal measurement resulted in  $P_s$  of  $\sim 4.6 \mu\text{C cm}^{-2}$ ,  $P_r$  of  $\sim 8.3 \mu\text{C cm}^{-2}$  and  $E_c$  of  $\sim 5.8 \text{ kV cm}^{-1}$  with the least leakage current. Further, based on the first-principles calculations discussed above, the  $P_s$  was predicted to be  $3.1 \mu\text{C cm}^{-2}$ , which agrees well with the experimental value. In addition to the polar FE form, this system exhibits two additional polymorphs: orthorhombic polar form ( $1P_O$ ,  $Pca2_1$ ) and tetragonal centrosymmetric form ( $I4_1/a$ ). However, the FE measurements on the pellet of  $1P_O$  did not result in any meaningful polarization switching loop. Moreover, the  $P_s$  value estimated using VASP is also nominal ( $0.2 \mu\text{C cm}^{-2}$ ). Therefore, the in-crystal ferroelectric measurements were carried out only in the case of the  $1P_T$  form.

#### 2.4.2 Optically controlled ferroelectrics

Other than electric fields and temperature, structural changes can also be triggered by stimuli, such as pressure, light, mechanical stress, and magnetic fields. Among them, photoirradiation is a noncontact, noninvasive, nondestructive and remotely controlled stimulus for optical control of FE polarization. Photoinduced geometrical isomerization, such as *cis-trans* or keto-enol transformation, can trigger a structural phase transition leading to the photochromism effect. Despite the discovery of photochromism as early as 1867, photoswitchable FEs are rare because of the impossibility of such structural changes in inorganic systems<sup>106</sup>.



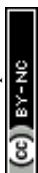
Fortunately, a special class of organic materials, such as azo *cis-trans*, Schiff base enol-keto, etc., undergo photoisomerization in their solid state<sup>107</sup>. The recent developments on optically controlled SCOF based on salicylideneaniline derivatives and azobenzene-based crystals have shed lights in the design of single-component photoswitchable FEs. Salicylideneaniline derivatives are a special class of organic Schiff base materials that undergo photochromism upon irradiation due to keto-enol tautomerism. This class of materials has recently been discovered as optically controlled SCOFs. Xiong *et al.* have reported multiple examples of salicylideneaniline derivatives to show photocontrol FE behaviour<sup>75,108–110</sup>. The first such example is 3,4,5-trifluoro-N-(3,5-di-tertbutyl salicylidene) aniline (TFTBSA)<sup>75</sup>, which crystallizes in two polymorphic forms, polar form with space group *Pna*2<sub>1</sub> (TFTBSA-1) and the centrosymmetric form with space group *P*2<sub>1</sub>/*n* (TFTBSA-2). The polymorphs have distinct molecular conformations, i.e. they are conformational polymorphs. The polar TFTBSA-1 crystals undergo a photochromic effect under the illumination of UV light (365nm), resulting in changes in colours from yellow to orange (Figure 10). In contrast, the centrosymmetric polymorph TFTBSA-2 does not change colours even after UV light is illuminated for longer. The structural switching due to photoirradiation in the case of the polar form has been confirmed through ultraviolet–visible (UV–vis) absorption, and infrared (IR) spectra are accompanied by a reversible structural photoisomerization between the enol form and the *trans*-keto form, which is impossible to be realized in inorganic FEs. In addition, under remote-control photoirradiation, the dielectric permittivity and *P*<sub>s</sub> can be switched reversibly by structural photoisomerization. In this case, the photoswitchable FE was achieved upon polarization switching through a structural phase transition triggered by photoisomerization. Furthermore, the authors also studied N-(3,5-ditert-butylsalicylidene) aniline, another salicylideneaniline derivative that crystallizes in the polar space group *Pna*2<sub>1</sub>. It was observed that despite fulfilling the structural criteria for FE, this derivative does not undergo polarization switching under the external *E*. This behaviour suggested that a compound must simultaneously have a polar point group and photoisomerization effect in this class of systems to achieve photo-controlled polarization switching. Moreover, in the quest for many more photocontrol FEs, they have come across multiple systems such (*R*)- and (*S*)-N-3,5-di-tert-butylsalicylidene-1-4-bromophenylethylamine, (*E*)-2-((phenyl amino)methyl)phenol, (*E*)-2-(((perfluorophenyl) imino)methyl) phenol<sup>108</sup>, salicylidene-2,3,4,5,6-pentafluoroaniline<sup>109</sup>, salicylidene aniline<sup>110</sup>, etc. Among these, (*R*)- and (*S*)-N-3,5-di-tert-butylsalicylidene-1-4-bromophenylethyl amine exhibit multichannel controlled multiferroic behaviour, where both photocontrol polarization and temperature-controlled ferroelastic domain has been observed. Further, such studies have been extended to the diarylethene<sup>111</sup>, fulgide<sup>112</sup>, muconic<sup>113</sup>, and azobenzene<sup>106</sup> based family of photochromic compounds.



**Figure 10:** Conversion of enol form to *trans*-keto form and vice versa upon UV light irradiation (up) and photochromic and non-photochromic behaviour of the polymorphic forms TFTBSA-1 and TFTBSA-2, respectively. *tBu* = *tert*-butyl. Reprinted (used in part) with permission from reference<sup>75</sup>. Copyright 2021 American Chemical Society.

#### 2.4.3 Disk and wheel-like ferroelectrics

Disk or wheel-like molecules have recently been used to develop SCOFs, where polarization switching is achieved via molecular rotation<sup>114</sup>. Such FEs follow the principle of latent pseudosymmetry in their crystal structure, which leads to the achievement of the hypothetical PE phase. This allows easy polarization switching when the crystal experiences minimal deviation from a centrosymmetric structure. In such molecules, the rotational disorder can awaken the hidden symmetry of the whole crystal structure. Such FEs can be disk-type, where in-plane rotation of molecules leads to polarization switching, and wheel-type, where axial rotation can flip the molecular dipole, leading to polarization switching. Hexa-substituted benzene derivative and 1,2,3,4,5-pentamethyl-6-nitrobenzene (PMNB)<sup>114</sup> have been utilized for the development of disk-type FEs (Figure 6). In the case of PMNB, the nitro-group is an electron-withdrawing group, and the penta-methyl group is an electron-donating group that forms a polar molecule which crystallizes in space group *P*1. The molecules in the crystal structure are assembled in a sheet parallel to (11  $\bar{1}$ ), where each molecule is located on the pseudo-inversion symmetry. This leads to an easy *P*<sub>s</sub> switching simply by the molecular disk rotation. Although the PMNB crystal exhibits a large *P*<sub>r</sub> of ~6  $\mu\text{C cm}^{-2}$ , it experiences a very high *E*<sub>c</sub> (~100  $\text{kV cm}^{-1}$ ) at RT and a frequency of 1 Hz along the *c*-axis. In the case of wheel-type FEs, the axially asymmetric substituents are used as a wheel or rotary knob. The molecules are designed in tetragonal-shaped substituents –RXYZ or triangular-shaped substituents –RXY fashion where a large dipole moment can be induced by the introduction of an electron-donating atom or group X and electron-withdrawing atom or group Y. Recently, 2-(methylsulfonyl)-malonamide (MSMA)<sup>114</sup> has been reported as wheel-type FE (Figure 6). MSMA experiences a net polarization of 7.3  $\mu\text{C cm}^{-2}$ , *P*<sub>r</sub> of 3.8  $\mu\text{C cm}^{-2}$ , and a high *E*<sub>c</sub> of ~100–160  $\text{kV cm}^{-1}$  at 3–30 Hz. Both PMNB and MSMA crystals exhibit a nearly temperature-independent relative permittivity at near RT, indicating the absence of a *T*<sub>c</sub>.



#### 2.4.4 Homochiral ferroelectrics

Combining quasi-spherical theory with homochiral principle dramatically improves the probability of discovering new SCOFs<sup>65, 115</sup> (Figure 6). The quasi-spherical principle assists in discovering material with phase transition behaviour of order-disorder type. Further, as discussed above, the homochirality enhances the possibility of a molecule crystallizing in chiral point groups. Recently Xiong *et al.* exploited this concept to achieve quasi-spherical homochiral FE crystals (*R*)- and (*S*)-3-quinuclidinol by introducing –OH group to the quinuclidinol molecule (Figure 11)<sup>115</sup>. While the (*R*)- and (*S*)-forms crystallize in enantiomorphic space groups  $P6_1$  and  $P6_5$ , respectively, the (*rac*)-3-quinuclidinol crystallizes in the centrosymmetric space group  $P2_1/n$ . Both the polar forms undergo FE to PE phase transition at ~400K via an order-disorder mechanism due to globular molecular structure with 622F6 type FE transition. They exhibit uniaxial characteristics with  $P_s$  of ~7  $\mu\text{C cm}^{-2}$  along O–H...N hydrogen bond propagating along the *c*-axis. The uniaxial characteristics of 3-quinuclidinol limit their practical utility. Thus, the quest for FE candidates with multiaxial characteristics based on the quasi-spherical principle led to the discovery of 2-(hydroxymethyl)-2-nitro-1,3-propanediol<sup>116</sup>, in which a six-fold vertex domain structure was observed for the first time. In this multiaxial FE, the transition follows  $m3mF1$  Aizu notation, and the change in the symmetry elements resulted in 24 equivalent FE axes and 48 crystallographic equivalent polarization directions, the highest among all ferroelectrics. Moreover, the  $P_s$  of 8.4  $\mu\text{C cm}^{-2}$ , observed in this system, is comparatively higher than that of the well-known FE polymer PVDF. Furthermore, recently discovered, homochiral SCOFs, (*R*)-10-camphorsulfonylimine and (*S*)-10-camphorsulfonylimine<sup>117</sup>, under the guidance of FE chiral chemistry have been reported to exhibit highest  $T_c$  of 429 K among the SCOFs.

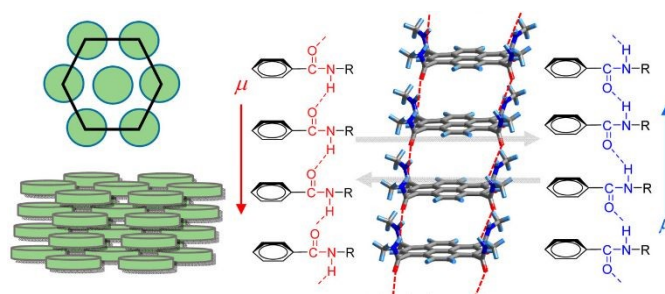


**Figure 11:** Scheme representing the introduction of homochirality in 3-quinuclidinol. Reprinted (used in part) with permission from reference<sup>118</sup>. Copyright 2020 American Chemical Society.

#### 2.4.5 Discotic type ferroelectrics

Flexible supramolecular assemblies with directional amide-type N–H...O hydrogen bonds form molecular columns, leading to discotic-type materials with large macroscopic dipoles along the column axis. Based on this design principle, discotic liquid crystalline (DLC) materials with benzene-1,3,5-tricarboxamide (BTA) core, forming hexagonal 1D columnar phase with varying lengths of alkyl chains containing C6, C10, and C18 carbons (Figure 6) were used to study dielectric relaxation and ferroelectric/polar switching<sup>119,120</sup>. While the C18-based DLC yields a significant  $P_r$  ( $=P_s$ ) value of 1.3 – 1.8  $\mu\text{C cm}^{-2}$  and  $E_c$  value of 20–30  $\text{kV cm}^{-1}$  in its thin film<sup>120</sup>, the temperature-dependent experiments in crystals of C10 and C18-based DLC result in extrinsic polar switching with  $P_s$  and  $E_c$  values of 1 – 2  $\mu\text{C cm}^{-2}$  and

20–30  $\text{kV cm}^{-1}$ , respectively. Further, Akutagawa *et al.* reported that the benzenecarboxamide derivatives bearing four –CONHC<sub>14</sub>H<sub>29</sub> chains at the pyrene core form  $\pi$ -stacked column through intermolecular N–H...O=C hydrogen bond and display FE switching.<sup>121</sup> In the case of the hexacolumnar DLC phase above 295K, the inversion of the dipole moment due to the conformation change of the hydrogen bond along the  $\pi$ -stacking column under  $E$  leads to the generation of FE switching (Figure 12). Further, they used non- $\pi$ -planar helicene cores bearing two –CONHC<sub>14</sub>H<sub>29</sub> chains (Figure 6), forming discotic molecular assemblies with 2D intermolecular interactions to achieve superior FE properties and enhance the practical utility of such systems<sup>122</sup>. They have successfully measured excellent FE properties ( $P_r$  and  $E_c$  values of 10.7  $\mu\text{C cm}^{-2}$  and 10.8  $\text{kV cm}^{-1}$ , respectively) in *rac*-N,N'-Di(tetradecyl)-1,12-dimethyl-N<sup>5</sup>,N<sup>8</sup>-ditetradecylbenzo[*c*]phenanthrene-5,8-dicarboxamide which forms high-density and well-oriented ferroelectric domains in a 2D hydrogen-bonding network layer.



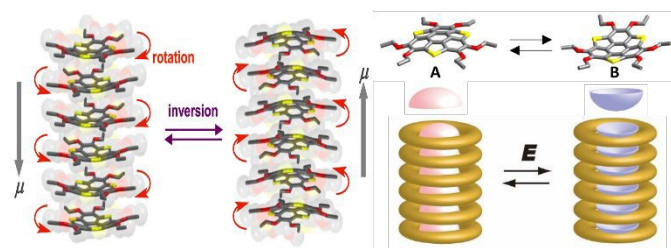
**Figure 12:** Schematic illustration of discotic hexacolumnar mesophase (left) and H-bonded 1D columnar structure with dipole ( $\mu$ ) inversion through the collective conformational change of the chain along the  $\pi$ -stacking direction (right). Reprinted (used in part) with permission from reference<sup>121</sup>. Copyright 2015 American Chemical Society.

#### 2.4.6 Bowl-to-bowl inversion type ferroelectrics

Supramolecular assemblies based on organic molecules with a bowl-shaped  $\pi$ -aromatic core that undergoes bowl-to-bowl inversion and skew rotation generate ferroelectric dipole relaxation<sup>123,124</sup>. Organic ferroelectric material based on hexaalkoxy trithiasumanenes (CnSS), which feature invertible bowl-shaped  $\pi$ -aromatic cores, is one such finding<sup>125</sup>. The study reported that the bowl-shaped aromatic cores of these CnSS molecules were stacked one-dimensionally in a columnar arrangement (Figure 13). Such an alignment resulted in a dipole moment from the stacked bowl-shaped framework, leading to remanent polarization in the crystals. Generally, the bowl inversion barrier magnitude depends on the bowl depth of the curved  $\pi$  frameworks. The shallow depth of Trithiasumanene (SS) led to a low bowl inversion barrier. Moreover, introducing large heteroatoms such as sulphur and selenium in curved  $\pi$ -molecules impacted the dynamic and electronic properties of these heterasumanenes. Six long alkoxy chains (–OC<sub>n</sub>H<sub>2n+1</sub>) were introduced into the SS  $\pi$ -core to enhance the motional freedom and facilitate bowl inversion. The above room temperature FE reversal polarizations of the CnSS series were confirmed via differential scanning calorimetry and  $P$ – $E$  hysteresis loop measurements at 336 – 383 K.







**Figure 13:** Schematic dipole inversion mechanism coupled with inversion of bowl-shaped (from A (up) to B (down)) conformation and skew rotation and a schematic model of the bowl-to-bowl structural inversion upon application of an electric field  $E$ , where molten  $-\text{OC}_n\text{H}_{2n+1}$  chains surround the polar bowl column. The grey arrows indicate the dipole moment direction in each system (adapted and modified from reference<sup>125</sup>).

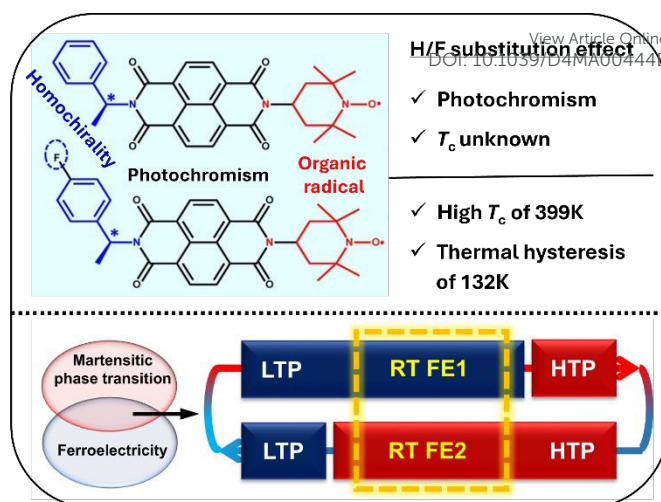
#### 2.4.7 Organic radical and martensitic ferroelectrics

Organic radicals, 2,2,6,6-Tetramethylpiperidine-N-oxyl (TEMPO)<sup>126</sup> and  $[(\text{NH}_3\text{-TEMPO})\text{([18]crown-6)}](\text{ReO}_4)$ <sup>127</sup> are well-known single- and multi-component FE materials, respectively. However, they display either a low or no phase transition temperature ( $T_c$ ). In this context, it is worth recalling that the  $T_c$  usually remains inaccessible for SCOFs. Nonetheless, the martensitic phase transition, which occurs in solids due to lattice strain,<sup>128,129</sup> has been reported in organic FEs, such as guanidinium nitrate<sup>130</sup> and diisobutylammonium bromide<sup>131</sup>. Based on these design elements for SCOFs, Zhang *et al.*<sup>132</sup> studied a series of asymmetric 1,4,5,8-naphthalenediimide (NDI) derivatives that form lamellar structures via  $\pi\cdots\pi$  interactions, a key strategy for introducing martensitic phase transition in organic compounds. Further, to induce ferroelectricity in NDI derivatives, they attached homochiral  $\alpha$ -methyl benzylamine on one side and  $\text{NH}_2\text{-TEMPO}$  on the other to construct organic radical ferroelectrics (Figure 6). Subsequently, they synthesized a series of FE compounds (N-[(S)- and (R)-1-phenylethyl]-N'-[1-oxyl-2,2,6,6-tetramethyl-4-piperidinyl]-1,4,5,8-naphthalenediimide (S- and R-H) and their 4-fluorophenyl analogues (S- and R-F) (Figure 14). The study revealed that the derivatives adopt a low-symmetric polar structure and display FE properties. Further, the study demonstrated that upon H/F substitution, the fluorinated compounds experience reversible FE and martensitic transitions at 399 K, with a significant thermal hysteresis of 132 K.

Additionally, biodegradable ferroelectricity is reported in molecular crystals of 2,2,3,3,4,4-hexafluoropentane-1,5-diol (HFPD), a single-component system that exhibits large piezoelectric response<sup>133</sup>. The study demonstrated that the crystals can flexibly be made into thin, strain-sensing films. Moreover, it exhibited promising biocompatibility with biological cells and desirable biodegradation and biosafety in physiological environments. Thus, it was envisaged that the material may find applications in transient implantable electrochemical devices.

#### 2.5. Advance methodologies to study H-bonded FEs

Conventionally, FE characterization techniques include X-ray diffraction for phase and structural analysis,  $P$ - $E$  loop measurements



**Figure 14:** Effects of H/F substitution in NDI derivatives (top) and the combination of martensitic phase transition and ferroelectricity in S-/R-F NDI derivatives with two different ferroelectric phases at room temperature (bottom). LTP and HTP represent the low-temperature phase and high-temperature phase, respectively. RT FE1 and RT FE2 represent different ferroelectric phases at room temperature (adapted and modified from reference<sup>132</sup>).

to determine the macroscopic polarization and switching behaviour, and PFM experiment to visualize the domain structure and switching pattern based on amplitude and phase measurements to know the local FE behaviour. For H-bonded FEs, in addition to the conventional methodologies, advanced techniques such as neutron diffraction<sup>134</sup> and high-resolution single-crystal X-ray diffraction experiments can be employed to understand better the FE mechanisms, especially for the order-disorder and PTM-type FEs. Charge density analysis using high-resolution X-ray and neutron diffraction data allows deriving accurate topological properties of electron densities. While neutron diffraction pinpoints the position of hydrogen atoms, high-resolution X-ray diffraction provides an opportunity to measure the degree of polarization of electron densities, which are crucial for PTM-type FEs. Moreover, neutron data allows the accurate determination of the atomic displacement parameters and leads to the efficient resolution of atomic disorders. In addition,  $E$ -dependent X-ray and neutron diffraction experiments may help monitor the dynamics of the proton due to FE switching and the corresponding  $\pi$ -electron delocalization. Recently, Fancher *et al.* performed the  $E$ -dependent neutron diffraction study on KDP crystals, a H-bonded FE<sup>135</sup>. The study demonstrates that the macroscopic polarization reorientation is instigated by a collective change in the H-bond, leading to the reversal of spontaneous dipole. Additionally, the decrease of secondary extinction near the  $E_c$  suggested that the inversion of macroscopic polarization is due to the nucleation and the successive growth of the new domains. Moreover, Kobayashi *et al.* observed electric field-induced antipolar-polar transition, along with colossal electrostriction in 2-trifluoromethyl-naphthimidazole crystal by performing an  $E$ -dependent synchrotron X-ray diffraction experiment<sup>136</sup>. These few studies based on advanced techniques provided crucial insights towards the improved understanding of the FE mechanisms in H-bonded FEs.





### 3. Negative Thermal Expansion

#### 3.1 Mechanism of NTE

The NTE phenomenon is generally noticed in inorganic systems. Most isotropic NTE is reported in open framework structures such as  $\text{AM}_2\text{O}_8$ <sup>80</sup>,  $\text{AM}_2\text{O}_7$ <sup>81</sup>,  $\text{A}_2\text{M}_3\text{O}_{12}$ <sup>82</sup>, zeolites<sup>83–85</sup> and ruthenium-based fluorides<sup>86–88</sup>. These materials experience the NTE effect due to the transverse vibration of the bridging atom, which includes oxygen and fluorine. The other reports on the mechanisms driving the NTE phenomenon are magnetic transition<sup>89</sup>, charge transfer<sup>90,91</sup> and spontaneous volume ferro-electrostriction<sup>92,93</sup>, which are mostly prominent in inorganic functional materials. The possible NTE mechanisms in inorganic solids and molecular materials are summarized in the following sections.

##### 3.1.1 Transverse vibration

Transverse vibration has been one of the most common mechanisms in inorganic solids (**Figure 15**). A few examples are the NZP family<sup>29</sup>,  $\text{ZrW}_2\text{O}_8$ <sup>30</sup>,  $\text{AM}_2\text{O}_7$ <sup>33</sup> (A = Ti, Zr, Hf, M = P, V),  $\text{Sc}_2(\text{WO}_4)_3$ <sup>137</sup>, zeolites<sup>35–37</sup>,  $\text{Zr}_{0.4}\text{Sn}_{0.6}\text{Mo}_2\text{O}_8$ <sup>32</sup>,  $\text{ScF}_3$ <sup>38</sup>, and cyanide-containing molecular materials such as  $\text{Fe}[\text{Co}(\text{CN})_6]$ <sup>138</sup>,  $\text{Zn}(\text{CN})_2$ <sup>139</sup>,  $\text{Cd}(\text{CN})_2$ <sup>140</sup>,  $\text{N}(\text{CH}_3)_4\text{CuZn}(\text{CN})_4$ <sup>141</sup>,  $\text{MPT}(\text{CN})_6$ <sup>142</sup> (M = Mn, Fe, Co, Ni, Cu, Zn, Cd), and  $\text{LnCo}(\text{CN})_6$ <sup>143</sup> (Ln = La, Pr, Sm, Ho, Lu, Y). These materials form a network of corner-shared polyhedrons due to bridging atoms such as oxygen and fluorine between the metal ions. This central linking atom in the chain M1–O–M2 oscillates perpendicular to the chain upon heating, resulting in contraction of the distance between M1...M2 (red colour atoms in **Figure 15**), and the material exhibits NTE behaviour macroscopically. An alternative way to derive this mechanism is rigid unit modes (RUMs), whereby polyhedral can be treated as rigid units due to strong bonding between M–O. The dynamic rocking motion of the rigid polyhedrons causes NTE due to the temperature rise.

Williams *et al.*<sup>139</sup> used cyanide molecule as the bridging agent in  $\text{Zn}(\text{CN})_2$  to enhance the NTE ( $\alpha_L = -18 \text{ MK}^{-1}$ ) below 180 K compared to inorganic oxide such as  $\text{ZrW}_2\text{O}_8$  with single-atom-supported RUM structures. The enhanced NTE effect in the case of diatom-supported RUMs is due to larger degrees of freedom. RUMs supported by a single atom are face-constrained, where an adjacent polyhedron influences the rotation of a polyhedron. However, for the diatom-supported RUMs, the polyhedron can rotate independently. Two neighbouring squares can rotate along the same or opposite directions, contributing to NTE.

##### 3.1.2 Host-guest interaction and geometrical flexibility

The concept of host-guest chemistry was first established by Donald J. Cram<sup>144</sup> in 1974 to illustrate the chemistry of a multi-component system comprised of two or more molecules or ions held together by noncovalent interactions. Because the NTE effect is highly prone to crystal lattice dynamics, the introduction of guest molecules in the voids of crystal lattice strongly influences the

thermo-responsive behaviour of the molecule. Furthermore, molecular framework deformation due to geometrical flexibility under external stimuli such as temperature, pressure, guest molecule, light, and irradiation leads to outstanding stimuli-responsive behaviours. Guest molecule plays an interesting role in modulating thermal expansion properties in some host-guest complexes. In 1990, White *et al.* performed the first thermal expansion study on a clathrate of Dianin's compound, the host-guest system<sup>145</sup>. In this case, the high TE, particularly along the *a*-axis, was achieved due to an increase in the anharmonicity of the guest molecule at high temperature compared to the guest-free form. Supramolecular interactions between the host-guest compounds lead to high thermal expansion as weakly interacted atoms of the guest molecule tend to vibrate more at high temperatures. Monohydrates of dipeptide tryptophylglycine exhibit NTE due to the presence of water molecules in its crystal structure<sup>146</sup>.

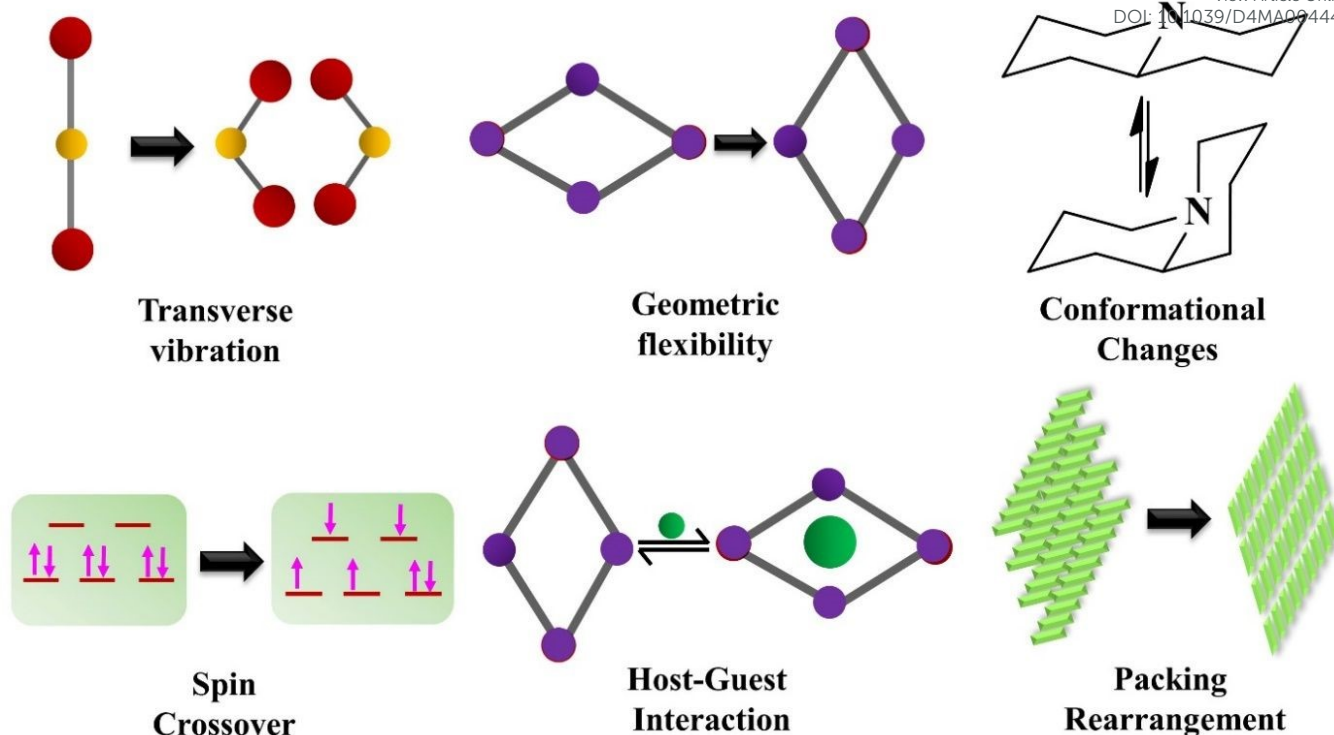
Geometrical flexibility is commonly observed in the case of MOF, which is prepared by extended ligands, resulting in a larger void within the crystal lattice and leading to a flexible structure (**Figure 15**). One of the well-known examples of flexible MOF is MIL-53<sup>147</sup>, which undergoes large pore ( $l_p$ ) and narrow pore ( $n_p$ ) transformation upon inclusion of guest molecule. This effect is called the breathing effect due to chemical stimuli. Further, geometrical flexibility due to temperature as a stimulus has been observed in guest-free aluminium-based MIL-53, where  $n_p$  to  $l_p$  transformation happens above 300K, and upon cooling, it undergoes a reversible process with a large thermal hysteresis. Furthermore, it has been observed that materials with fence-like structures undergo geometrical flexibility due to scissor-like motion<sup>148</sup>. Perdeuterated methanol monohydrate<sup>149</sup> is a simple molecular crystal which shows remarkable thermo-responsive behaviour with colossal anisotropic thermal expansion with  $\alpha_b = 462(4) \text{ MK}^{-1}$ ,  $\alpha_c = 93(4) \text{ MK}^{-1}$ , and  $\alpha_a = -61(4) \text{ MK}^{-1}$ , along the *b*, *c*, and *a*-axes, respectively, at 155 K due to scissor-like motion. With the rise in temperature, the dynamic molecular motion of the scissor-jack-like structures leads to contractions in the orthogonal direction and an expansion in the other direction.

Two salts of imidazole, imidazolium terephthalate (IMD-TPH) and imidazolium 4-hydroxybenzoate (IMD-HBC), were reported to have uniaxial NTE and biaxial PTE characteristics<sup>150,151</sup>. The unusual behavior of the salts was elucidated based on their fencing and scissor-like motion of the 2D H-bonded network, respectively. Upon changing the imidazole with 4-methylimidazole, salt 4-methylimidazolium-4-hydroxybenzoate (4-MIMD-HBC) exhibited unusual biaxial NTE owing to the scissor-like motion with the variation of temperature<sup>152</sup>.

##### 3.1.3 Molecular packing rearrangements and conformations

Purely organic molecular crystals comprising weak noncovalent interaction<sup>153–155</sup> facilitate the packing rearrangement and conformational change, causing anisotropic thermo-responsive effects upon thermal stimuli (**Figure 15**). These materials are





**Figure 15:** Various mechanisms leading to negative thermal expansion behaviour in molecular materials.

categorized as soft materials, as the molecules in these materials are held together via weak supramolecular interactions. The molecular packing rearrangement leading to exceptionally large reversible uniaxial PTE ( $156 \text{ MK}^{-1} < \alpha_L < 515 \text{ MK}^{-1}$ ) and biaxial NTE ( $-32 \text{ MK}^{-1} < \alpha_L < -85 \text{ MK}^{-1}$  and  $-48 \text{ MK}^{-1} < \alpha_L < 204 \text{ MK}^{-1}$ ) was observed by Das and co-workers<sup>156</sup> in the case of dumbbell-shaped (S, S)-octa-3,5-diyne-2,7-diol. The colossal anisotropic PTE and NTE were confirmed based on the variable temperature (VT) SCXRD study. The study demonstrated that decreasing the temperature from 330 to 240 K results in a gradual shift in the molecular orientation, leading to the contraction along the *a*-axis and expansions along the *b*- and *c*-axes.

In organic molecular crystals, the conformational changes and packing rearrangements lead to the structures crystallizing in polymorphic forms, which often display distinct thermo-responsive effects, as was the case for the dimorphic 4,4'-methylenebis(2,6-dimethylaniline)<sup>157</sup>. The two forms crystallize in triclinic and monoclinic crystal systems and exhibit distinct thermal properties due to the striking molecular packing differences. While one polymorph experiences regular PTE, the other undergoes anisotropic NTE. Panda *et al.*<sup>158</sup> observed distinct thermo-responsive behaviour in two packing polymorphs of thermo-salient crystal of (phenylazopenyl) palladium hexafluoroacetylacetonate. The  $\alpha$ -form with head-to-head molecular stacking exhibits uniaxial NTE of  $-79.2 \text{ MK}^{-1}$  along the *c*-axis and biaxial PTE of  $260.4$  and  $39.4 \text{ MK}^{-1}$  along the *a*- and *b*-axes. However, the  $\beta$ -form with head-to-tail molecular stacking shows the usual PTE.

### 3.1.4 Spin crossover

Molecular materials comprised of  $3d^4$ - $3d^7$  transition metal, due to their electronic configuration, can undergo spin-crossover (SCO) phenomenon involving the switching process between two molecular spin states (so-called high-spin, HS and low-spin, LS configurations) under external stimuli<sup>159,160</sup> (Figure 15). The stimuli such as temperature, pressure, light irradiation, magnetic field, and guest molecules causing the SCO effect led to spectacular changes in numerous physical properties such as magnetic optical and dielectric properties. The change in the  $3d$  electronic configuration due to SCO switching causes significant changes in the bond length and the corresponding molecular volume. For example, Fe (II) complexes<sup>160</sup> with octahedral geometry have been the most studied SCO systems, where  $3d$  orbital are split into two sub-orbitals,  $t_{2g}$ -triply degenerate low-energy orbital and  $e_g$ -doubly degenerate higher energy orbital. Thus, in the LS state, six d-orbital electrons will occupy the three degenerate  $t_{2g}$  orbital having the electronic configuration of ( $^6t_{2g}, ^0e_g$ ), while in the HS state, it adopts the ( $^4t_{2g}, ^2e_g$ ) configuration. The electronic effect of SCO leading to electron-phonon coupling causing remarkable crystal lattice changes was exploited by Kepert *et al.*<sup>161</sup> in designing flexible Fe (II) complex of Fe [Au (CN)<sub>2</sub>]<sub>2</sub>, which undergoes SCO at 221-226 K. Further, the change in lattice deformation was controlled from abrupt to gradual by introducing Ni into the complex, which is a non-spin crossover metal. Monotonous thermomechanical motion is clearly exhibited by the [Fe<sub>0.84</sub>Ni<sub>0.16</sub>] composition, which shows extreme lattice flexing. The CTE of (Fe<sub>0.84</sub>Ni<sub>0.16</sub>) [Au (CN)<sub>2</sub>]<sub>2</sub> at 215 K are  $-3200$ ,  $5200$ , and  $1500 \text{ MK}^{-1}$  along the three axes, respectively, which are at least one order of magnitude greater than those reported for other materials.



### 3.2 NTE in single-component organic materials

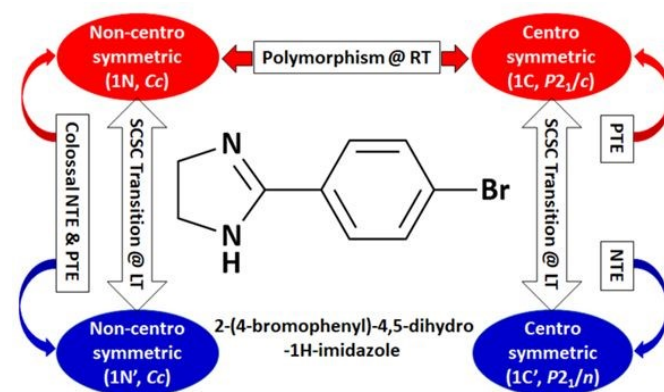
A recent study based on a literature survey highlighted that organic materials have received great attention in the area of NTE. The study revealed that about 30% of the molecular crystals show uniaxial NTE, whereas about 5% show biaxial NTE characteristic<sup>162</sup>. However, NTE characteristics in the case of purely organic systems are scarce. Recently, the development of organic materials with thermo-responsive properties<sup>7</sup> has been emphasized due to certain advantages such as flexibility, benign nature, and being environmentally friendly. The benign nature of the organic system allows it to adapt to physiological conditions. Thus, organic NTE materials can be utilized in dental filling. Additionally, molecular materials in their condensed phase are held together via weak interactions, comprising of H-bonds, coordination bonds, and weak van der Waals interactions such as  $\pi\cdots\pi$  interaction, halogen bonds, etc.<sup>153–155</sup>. Weak interactions are highly sensitive to the temperature. Therefore, a high PTE is expected for organic materials<sup>163</sup> in the direction of weak interactions<sup>164–166</sup>. Recently, the role of weak intermolecular interactions in controlling thermal expansion in molecular materials has been highlighted on a few occasions. The weak interactions provide enough flexibility due to a large degree of freedom, which results in some intriguing thermo-responsive and mechano-responsive behaviours such as NTE, breathing effect, and negative linear or area compressibility.

Switching of uniaxial PTE to NTE in two compounds, 1,4-bis(S-3-butyn-2-ol)-benzene and 2,6-bis(S-3-butyn-2-ol)-naphthalene with similar packing was reported in 2018<sup>166</sup>. The study claimed that the molecular packing arrangement and transverse vibration of some of the atoms led to this switching phenomenon. The report highlighted that the shortening of molecular length with the increase in temperature occurs due to the high thermal vibration of atoms. Very recently, Koshima *et al.* studied the switching of axial NTE to PTE in an enol form of N-3,5-di-tert-butylsalicylide-3-nitroaniline<sup>167</sup>. The material also exhibits photo actuation as it undergoes photoisomerization upon light irradiation.

#### 3.2.1 NTE in 4-substituted derivatives of imidazolines

In the course of the discovery of SCOFs<sup>61</sup>, some of us have explored the thermo-responsive behaviour of imidazoline derivatives (chloro, bromo, iodo, methyl and methoxy), which comprise an enamine unit (**Figure 7a**). The bromo derivative, 2-(4-bromophenyl)-4,5-dihydro-1H-imidazole<sup>168</sup>, exhibits dimorphism at RT (highlighted in red colour) to form non-centrosymmetric (1N, space group *Cc*) and centrosymmetric (1C, space group *P2<sub>1</sub>/c*) forms, as portrayed in **Figure 16**. While polymorphic form 1N (high temperature, HT phase) exhibits colossal NTE and PTE with axial switching, polymorph 1C shows only colossal PTE. Further, both forms undergo reversible iso-symmetric single-crystal-to-single-crystal (SCSC) phase transition at low temperature (LT) to form the LT phases (highlighted in blue colour) with space groups *Cc* (1N') and *P2<sub>1</sub>/n* (1C'), respectively. Interestingly, all four forms belong to the monoclinic crystal system, and three experience unusual thermal expansions. Like form 1N, phase 1N' (LT phase) also exhibits colossal NTE and PTE with axial switching. While 1N shows biaxial NTE of -

175.2 and -45.76 MK<sup>-1</sup> along the principal axes X<sub>1</sub> and X<sub>2</sub>, respectively, 1N' shows uniaxial NTE of -140.30 MK<sup>-1</sup> along the principal axis X<sub>1</sub>. Moreover, LT phase 1C' experiences a moderate biaxial NTE of 37.46 and -6.09 MK<sup>-1</sup> along principal axes X<sub>1</sub> and X<sub>2</sub>, respectively. The structure-property relationship study demonstrates that the interplay of the H-bonds and weak interactions leading to the scissor-jack-like molecular motion result in the anisotropic unusual thermal expansions over an extensive temperature range (~40 K to ~110 K) with record-wide thermal hysteresis. Further studies on halophenyl-substituted di-hydro imidazoles suggest that halogens play a crucial role in governing the unusual expansions in organic materials.



**Figure 16:** Polymorphism, phase transition, and unusual thermal expansion characteristics in halogen derivatives of 4-phenyl substituted imidazoline.<sup>168</sup>

Further, some of us reported polymorphism and hydration-driven thermo-responsive behaviour in 2-(4-(trifluoromethyl)-phenyl)-1H-phenanthro [9,10-d] imidazole (**Figure 4**)<sup>61</sup>. One of the three polymorphs, the orthorhombic form 1P<sub>O</sub>, undergoes anisotropic reversible switching from NTE (-55 MK<sup>-1</sup>) to PTE (25MK<sup>-1</sup>) along the cell edge *b* at below RT (~210 K). Moreover, while the non-centrosymmetric hydrated polymorph (1N<sub>H</sub>, space group *P1*) exhibits NTE in the temperature range of 100–280 K, the centrosymmetric hydrated polymorph (space group *P1*) shows the usual PTE behaviour. During heating, reversible biaxial NTE (-38 MK<sup>-1</sup> and -11 MK<sup>-1</sup>) and PTE (187 MK<sup>-1</sup>) characteristics are also noticed for 1N<sub>H</sub>. The mechanism of NTE was investigated via molecular packing analysis in the crystal lattice as a function of temperature. In the case of 1P<sub>O</sub>, with Z' = 2, the A (pink) and B (red) molecules stack along [010] (**Figure 17a**). The interplanar distances between the two A molecules (*d<sub>A</sub>*) and that between the two B molecules (*d<sub>B</sub>*) vary synchronously with the variation of temperature, leading to the contraction and the expansion of cell edge *b* and the decrease and increase of the interplanar distances *d<sub>A</sub>* and *d<sub>B</sub>* before and after the switching temperature (~210 K), respectively (**Figure 17b**). Along cell edge *b*, with the increase in temperature, the molecules come closer to each other (the distances *d<sub>A</sub>* and *d<sub>B</sub>* decrease) until ~210 K. Therefore, cell edge *b* experiences NTE. After that, it switches to PTE when the molecules go apart (distances *d<sub>A</sub>* and *d<sub>B</sub>* increase). This synergistic decrease and increase of *d<sub>A</sub>* and *d<sub>B</sub>* with temperature leads to the anisotropic expansion and switching from NTE ↔ PTE



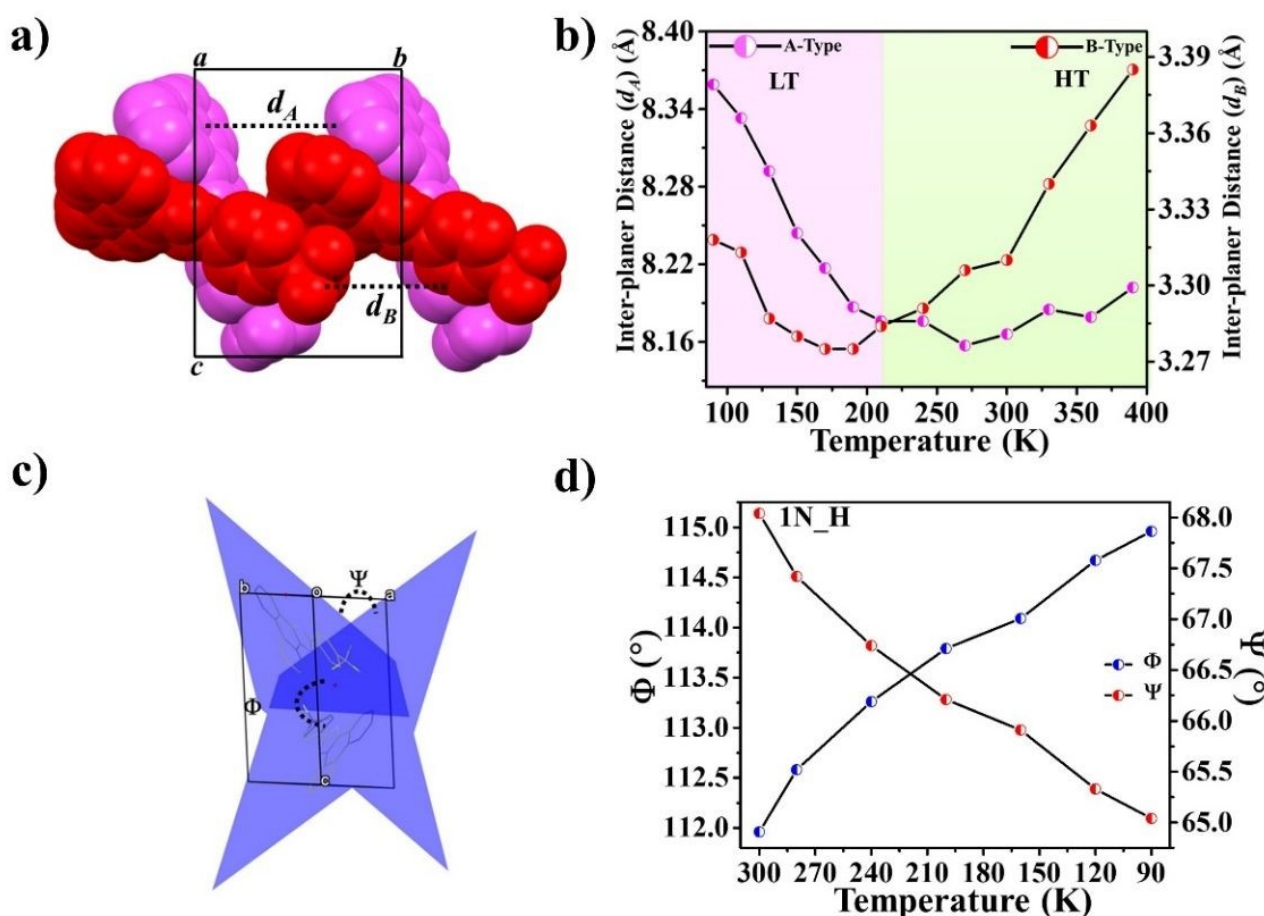


along cell edge  $b$  upon heating and cooling, respectively. For 1N\_H, the NTE and PTE arise due to the variation of the dihedral angles  $\Phi$  and  $\Psi$ , as shown in **Figures 17c** and **17d**. The molecular planes intersect in an open bookcase-like fashion and the opening and closing of the bookcase with temperature result in the expansion and contraction of the cell edges. While  $\Phi$  controls the variation of the cell edge  $c$ ,  $\Psi$  controls that of cell edges  $a$  and  $b$ . Upon heating, while the increase in  $\Phi$  leads to the expansion of cell edges  $a$  and  $b$ , the corresponding decrease in  $\Phi$  leads to the contraction of cell edge  $c$ .

### 3.3 NTE accompanied by order-disorder phase transformation

In the last few decades, the phase transition in molecular crystals has become an interesting topic of scientific investigation. The phase transitions are of particular interest as these studies provide the details of changes in molecular arrangements in the crystal structure

and can be compared through variable temperature X-ray diffraction (VTXRD) and differential scanning calorimetry studies. Using SCXRD and PXRD techniques, the phase transition can be monitored at variable temperatures (VT). VTXRD is an ideal technique for studying disordered systems as it allows distinguishing between static and dynamic disorders. The static disorder will remain the same at VT, while the dynamic disorder may disappear. However, a limited database is available for this type of phase transition, and most of the reports belong to inorganic and multicomponent organic materials. In 2014, Shang *et al.* reported order-disorder-induced NTE in two ammonium Mg-formate frameworks. Few materials from the  $ZrW_2O_8$  family are known to undergo order-disorder transition and show NTE.<sup>169</sup> Chen *et al.* reported H-bond driven order-disorder phase transition in an inorganic NLO material,  $[Ag(NH_3)_2]_2SO_4$ .<sup>170</sup> The order-disorder phase transition is reversible at 356 K and is associated with weak H-bonds in the crystal structure.



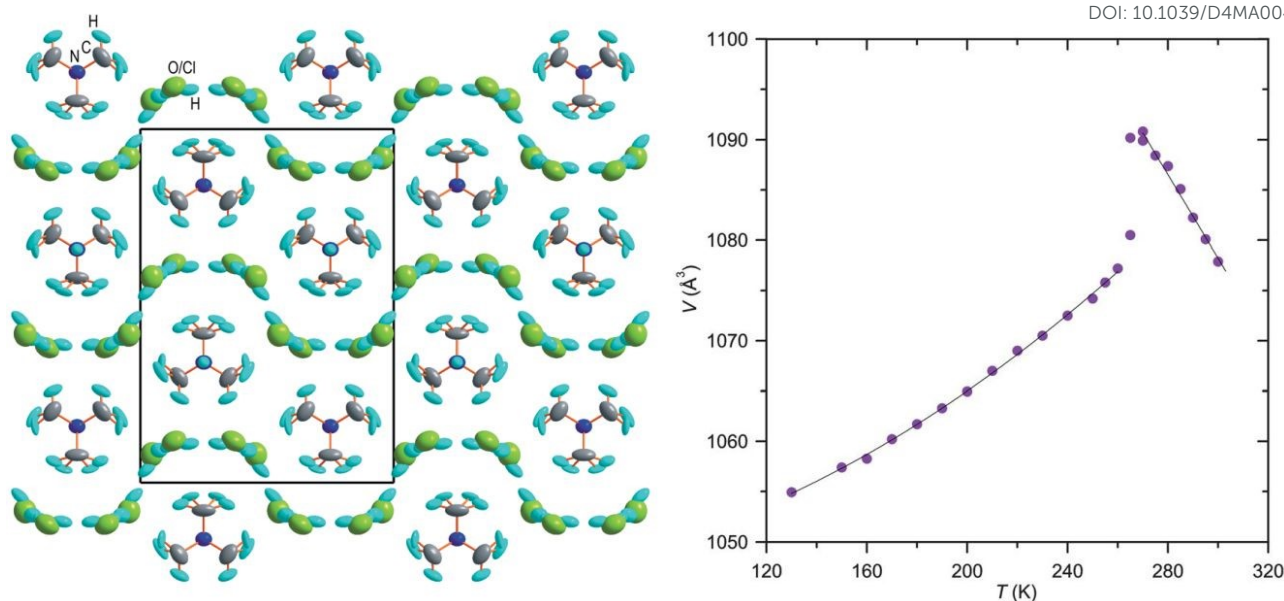
**Figure 17:** For 1P\_O (**Figure 4**), (a) stacking of molecule A (pink) and B (red) along [010] and (b) variation of inter-planar distances  $d_A$  and  $d_B$  with temperature. For 1N\_H (**Figure 4**), (c) dihedral angles  $\Phi$  and  $\Psi$  between the molecular planes passing through the adjacent molecules A and B, forming an open bookcase-like geometry while viewing down [110] and (d) the variation of  $\Phi$  and  $\Psi$  with temperature.<sup>61</sup>

The role of water molecules in inducing the NTE effect has been demonstrated on a few occasions. In the crystal structure of monohydrate dipeptide, tryptophylglycine ( $TrpGly \cdot H_2O$ ), the NTE effect in the helical direction ( $c$ -axis) occurs due to the increasing order of water molecules enclosed in the helices.<sup>124</sup> Szafranski reported that 1,3-diazabicyclo[2.2.2]-octane hydrochloride

trihydrate, ( $dabcoHCl \cdot 3H_2O$ ) undergoes phase transition at 260 K and displays very strong NTE along the  $a$  and  $c$ -axes ( $-529$  and  $33.3$   $MK^{-1}$ , respectively) at its HT phase. At 300 K, it experiences a large  $\alpha_V$  of  $-386$   $MK^{-1}$ , possibly the highest value in a molecular crystal, and the NTE occurs due to the rotation of rigid water molecules (**Figure 18**)<sup>50,146</sup>.



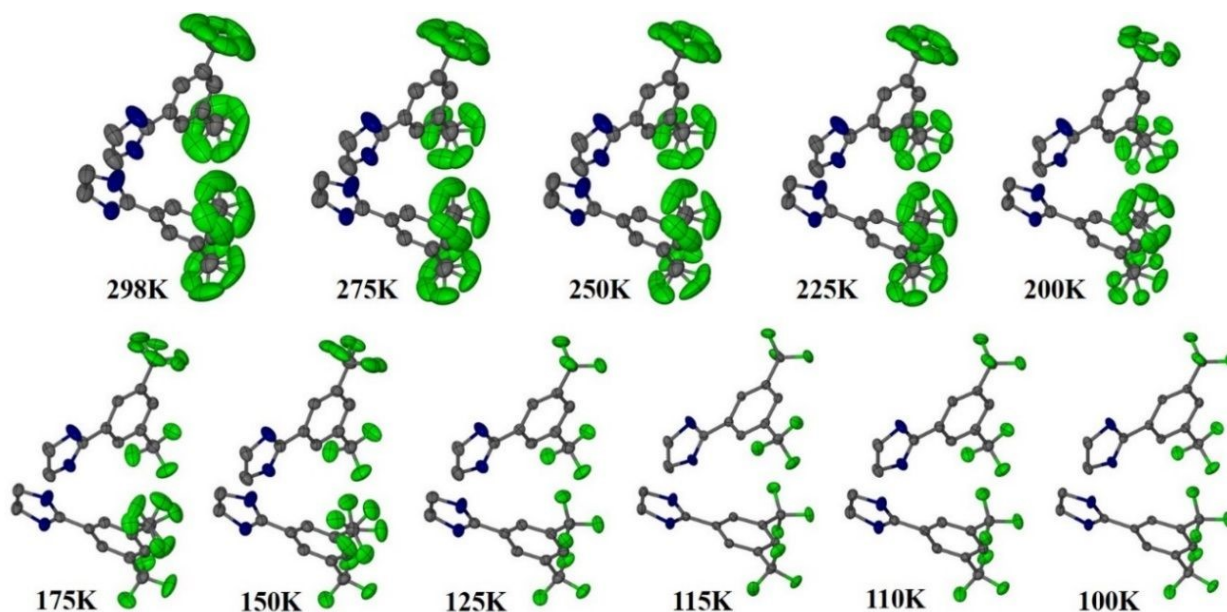




**Figure 18:** The crystal structure of the orthorhombic form of dabcoHCl.3H<sub>2</sub>O at 270K, viewed along [100], showing the H-bonded disordered H<sub>2</sub>O molecules and Cl<sup>-</sup> anions (Left). Temperature dependence of the unit-cell volume of dabcoHCl.3H<sub>2</sub>O (Right). Reprinted (used in part) from reference <sup>50</sup>.

Reports of NTE induced by order-disorder phase transition on single-component organic molecular crystals are extremely rare in the literature. Takahashi *et al.* reported two polymorphic forms of 2,4-dinitro anisole  $\alpha$  and  $\beta$ . In the case of  $\beta$  polymorph, the disordered in the nitro group becomes ordered at LT (261 K) and shows small biaxial NTE (-31 and -15 MK<sup>-1</sup>).<sup>171</sup> However, the  $\alpha$  polymorph exhibits usual PTE behaviour. Very recently, some of us reported thermal expansion behaviour of 2-(3,5)bis(trifluoromethyl)phenyl)-4,5-dihydro-1H-imidazole.<sup>172</sup> This lightweight material undergoes an iso-symmetric dynamic order-disorder phase transition at LT (<150 K) (**Figure 19**). Interestingly, the

material exhibits PTE, near ZTE, and NTE along its respective crystallographic axes (**Figure 20**) because it experiences a scissor-jack-like motion with the temperature variation. The N-H...N hydrogen bonds in the crystal structure act as a hinge to the scissor-jack and restrict the thermal expansion along the *b*-axis, leading to near ZTE behaviour. The HT (298 – 200 K) and the LT (125 – 100 K) phases display uniaxial NTE with CTE of -65 and -135 MK<sup>-1</sup> along X1 and biaxial PTE along X2 and X3, respectively. The material also exhibits a rare phenomenon of negative area expansion with CTE of -18 and -61 MK<sup>-1</sup> for the HT and LT phases, respectively.



**Figure 19:** Crystal structures of 2-(3,5)bis(trifluoromethyl)phenyl)-4,5-dihydro-1H-imidazole showing the transition from its disorder to order phase. H-atoms are omitted for clarity.<sup>172</sup>



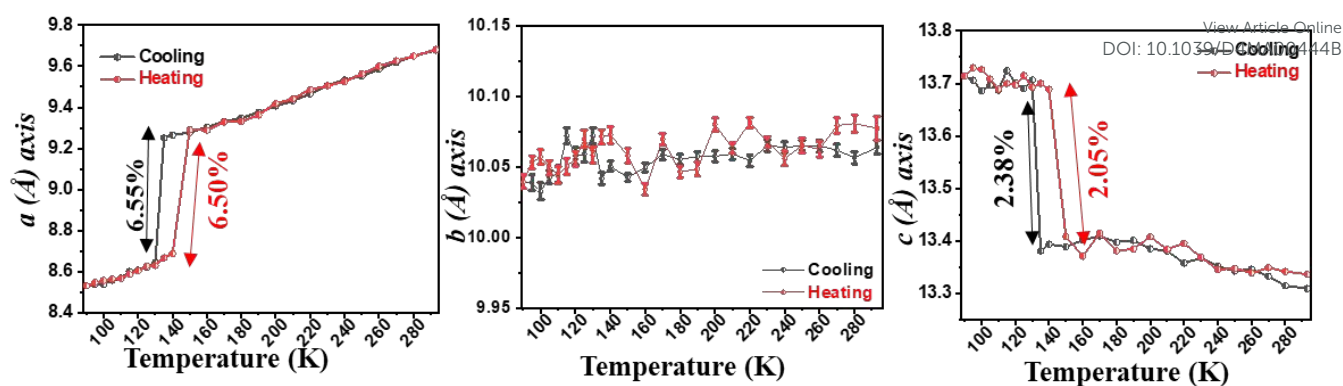


Figure 20: Variation of unit-cell lengths with temperature. The error bars highlighting the near ZTE behaviour are included for the *b*-axis.<sup>172</sup>

## 4. Conclusions and outlook

In this highlight, we provided an overview of the evolution of research on purely organic molecular crystals with multifunctional properties, FE and NTE, and the occurrence of polymorphism in such materials.

Initially, we covered the basics of FE and NTE properties and the origin of the coexistence of these intriguing phenomena. Subsequently, we introduced the polymorphism phenomenon and its role in governing molecular properties. Further, we highlighted the crucial role of crystallographic symmetries in designing the FE materials. Furthermore, we discussed the underlying mechanisms behind the unusual FE and NTE characteristics in these special classes of materials and the importance of crystallographic techniques in characterizing FE and NTE materials. Subsequently, we focused on purely organic materials, especially the single-component organic molecular crystals, and discussed a few important cases in detail. In many cases, we discussed the existence of polymorphic forms of the FE and NTE materials and the relationship of polymorphism with the FE and NTE characteristics. We also discussed the techniques typically used for characterising FE and NTE materials and the possibility of implementing various advanced experimental techniques and theoretical approaches to advance these key research areas in materials science.

Here, we emphasized the use of organic materials and the advantage of single-component systems and summarized various principles for designing SCOFs. It was evident from the highlight that there is plenty of scope for discovering new single-component FE and NTE materials. Although the co-existence of FE and NTE properties in an organic material is yet to be discovered, we discussed that the NTE characteristic has been observed in the two polymorphic forms of a SCOF, which was a step closure to combining the FE and NTE characteristics in a single system. We expect that this highlight will guide and motivate the researchers to design such materials for practical applications in the future.

Lastly, we highlighted that due to their benign nature, flexibility, lightweight, cost-effective production, easy processing, and least toxicity, organic materials with FE and unusual thermal expansion properties are preferred over inorganic materials for applications in flexible, wearable, biocompatible, reusable, economical medical devices. We envisage that this highlight will encourage scientists to

explore organic materials for preparing electronic devices with precise control of thermal expansion and superior memory and energy storage capacity, e.g., sensor-based non-expansive memory devices. Further research on multifunctional organic materials is expected to yield intelligent materials for environment-friendly applications, which may fulfil our societal needs.

## Author Contributions

PM has conceptualized and written the highlight. SD had written the original draft. LN contributed to the negative thermal expansion part. All the authors have contributed to editing the highlight.

## Conflicts of interest

There are no conflicts to declare.

## Acknowledgements

We thank Shiv Nadar University Institution of Eminence (SNIOE) for research facilities, infrastructure, funding, and research assistantship to SD and LN. This work is supported by research grants (CRG/2018/004694 and CRG/2023/008038) from SERB, Govt. of India.

## References

- 1 M. E. Lines and A. M. Glass, *Principles and Applications of Ferroelectrics and Related Materials*, Oxford University Press, Oxford, 2001.
- 2 J. F. Scott, *Ferroelectric Memories*, 2000.
- 3 J. F. Scott, *Science*, 2007, **315**, 954–959.
- 4 J. S. O. Evans, *J. Chem. Soc. Dalt. Trans.*, 1999, **0**, 3317–3326.
- 5 K. Takenaka, *Sci. Technol. Adv. Mater.*, 2012, **13**, 013001.
- 6 J. Chen, L. Hu, J. Deng and X. Xing, *Chem. Soc. Rev.*, 2015, **44**, 3522–3567.
- 7 Z. Liu, Q. Gao, J. Chen, J. Deng, K. Lin and X. Xing, *Chem. Commun.*, 2018, **54**, 5164–5176.
- 8 J. Bernstein, in *Strength from Weakness: Structural*



- Consequences of Weak Interactions in Molecules, Supermolecules, and Crystals*, Springer Netherlands, Dordrecht, 2002, pp. 247–260.
- 9 T. Miyazaki and H. Jin, *The Physics of Ferromagnetism*, Springer Berlin Heidelberg, Berlin, Heidelberg, 2012, vol. 158.
- 10 A. K. Tagantsev, L. E. Cross and J. Fousek, *Domains in Ferroic Crystals and Thin Films*, Springer New York, New York, NY, 2010.
- 11 Y. Noda, T. Yamada, K. Kobayashi, R. Kumai, S. Horiuchi, F. Kagawa and T. Hasegawa, *Adv. Mater.*, 2015, **27**, 6475–6481.
- 12 M. Fiebig, V. V. Pavlov and R. V. Pisarev, *J. Opt. Soc. Am. B*, 2005, **22**, 96.
- 13 N. E. Sherstyuk, E. D. Mishina, S. D. Lavrov, A. M. Buryakov, M. A. Marchenkova, A. S. Elshin and A. S. Sigov, *Ferroelectrics*, 2015, **477**, 29–46.
- 14 K. Uchino, *Ferroelectric Devices*, CRC Press, 2018.
- 15 G. Sebald, D. Guyomar and A. Agbossou, *Smart Mater. Struct.*, 2009, **18**, 125006.
- 16 J. Valasek, *Phys. Rev.*, 1920, **15**, 505–564.
- 17 J. Valasek, *Phys. Rev.*, 1921, **17**, 475–481.
- 18 G. Busch and P. Scherrer, *Naturwissenschaften*, 1935, **23**, 737–737.
- 19 G. H. Kwei, A. C. Lawson, S. J. L. Billinge and S. W. Cheong, *J. Phys. Chem.*, 1993, **97**, 2368–2377.
- 20 G. H. Haertling, *J. Am. Ceram. Soc.*, 1999, **82**, 797–818.
- 21 P. Theato, B. S. Sumerlin, R. K. O'Reilly and T. H. Epps, III, *Chem. Soc. Rev.*, 2013, **42**, 7055.
- 22 Y.-J. Kim and Y. T. Matsunaga, *J. Mater. Chem. B*, 2017, **5**, 4307–4321.
- 23 W. Miller, C. W. Smith, D. S. Mackenzie and K. E. Evans, *J. Mater. Sci.*, 2009, **44**, 5441–5451.
- 24 R. H. Baughman, *Science*, 2002, **297**, 787–792.
- 25 R. S. Krishnan, R. Srinivasan and S. Devanarayanan, *Thermal expansion of crystals: international series in the science of the solid state*, 2013.
- 26 M. J. Cliffe and A. L. Goodwin, *J. Appl. Crystallogr.*, 2012, **45**, 1321–1329.
- 27 R. M. Hazen and L. W. Finger, *Am. Sci.*, 1984, **72**, 143–150.
- 28 X. Yang, X. Cheng, X. Yan, J. Yang, T. Fu and J. Qiu, *Compos. Sci. Technol.*, 2007, **67**, 1167–1171.
- 29 R. Roy, D. K. Agrawal and H. A. McKinstry, *Annu. Rev. Mater. Sci.*, 1989, **19**, 59–81.
- 30 T. A. Mary, J. S. O. Evans, T. Vogt and A. W. Sleight, *Science*, 1996, **272**, 90–92.
- 31 K. Röttger, A. Endriss, J. Ihringer, S. Doyle and W. F. Kuhs, *Acta Crystallogr. Sect. B Struct. Sci.*, 1994, **50**, 644–648.
- 32 S. E. Tallentire, F. Child, I. Fall, L. Vella-Zarb, I. R. Evans, M. G. Tucker, D. A. Keen, C. Wilson and J. S. O. Evans, *J. Am. Chem. Soc.*, 2013, **135**, 12849–12856.
- 33 V. Korthuis, N. Khosrovani, A. W. Sleight, N. Roberts, R. Dupree and W. W. J. Warren, *Chem. Mater.*, 1995, **7**, 412–417.
- 34 J. S. O. Evans, T. A. Mary and A. W. Sleight, *J. Solid State Chem.*, 1997, **133**, 580–583.
- 35 D. A. Woodcock, P. Lightfoot, L. A. Villaescusa, M.-J. Díaz-Cabañas, M. A. Cambor and D. Engberg, *Chem. Mater.*, 1999, **11**, 2508–2514. DOI: 10.1039/D4MA00444B
- 36 D. A. Woodcock, P. Lightfoot, P. A. Wright, L. A. Villaescusa and M. A. Cambor, *J. Mater. Chem.*, 1999, **9**, 349–351.
- 37 P. Lightfoot, D. A. Woodcock, M. J. Maple, L. A. Villaescusa and P. A. Wright, *J. Mater. Chem.*, 2001, **11**, 212–216.
- 38 B. K. Greve, K. L. Martin, P. L. Lee, P. J. Chupas, K. W. Chapman and A. P. Wilkinson, *J. Am. Chem. Soc.*, 2010, **132**, 15496–15498.
- 39 L. Hu, J. Chen, A. Sanson, H. Wu, C. Guglieri Rodriguez, L. Olivi, Y. Ren, L. Fan, J. Deng and X. Xing, *J. Am. Chem. Soc.*, 2016, **138**, 8320–8323.
- 40 L. Hu, J. Chen, J. Xu, N. Wang, F. Han, Y. Ren, Z. Pan, Y. Rong, R. Huang, J. Deng, L. Li and X. Xing, *J. Am. Chem. Soc.*, 2016, **138**, 14530–14533.
- 41 M. van Schilfgaarde, I. A. Abrikosov and B. Johansson, *Nature*, 1999, **400**, 46–49.
- 42 Y. W. Long, N. Hayashi, T. Saito, M. Azuma, S. Muranaka and Y. Shimakawa, *Nature*, 2009, **458**, 60–63.
- 43 M. Azuma, W. Chen, H. Seki, M. Czapski, S. Olga, K. Oka, M. Mizumaki, T. Watanuki, N. Ishimatsu, N. Kawamura, S. Ishiwata, M. G. Tucker, Y. Shimakawa and J. P. Attfield, *Nat. Commun.*, 2011, **2**, 347.
- 44 J. Chen, K. Nittala, J. S. Forrester, J. L. Jones, J. Deng, R. Yu and X. Xing, *J. Am. Chem. Soc.*, 2011, **133**, 11114–11117.
- 45 J. Chen, F. Wang, Q. Huang, L. Hu, X. Song, J. Deng, R. Yu and X. Xing, *Sci. Rep.*, 2013, **3**, 2458.
- 46 Z. Pan, Y.-W. Fang, T. Nishikubo, L. Hu, S. Kawaguchi and M. Azuma, *Chem. Mater.*, 2022, **34**, 2798–2803.
- 47 Q. Li, K. Lin, Z. Liu, L. Hu, Y. Cao, J. Chen and X. Xing, *Chem. Rev.*, 2022, **122**, 8438–8486.
- 48 P. Hu, J. Chen, X. Sun, J. Deng, X. Chen, R. Yu, L. Qiao and X. Xing, *J. Mater. Chem.*, 2009, **19**, 1648.
- 49 S. Li, K. Takahashi, R.-K. Huang, C. Xue, K. Kokado, N. Hoshino, T. Akutagawa, S. Nishihara and T. Nakamura, *Chem. Mater.*, 2023, **35**, 2421–2428.
- 50 M. Szafranski, *J. Mater. Chem. C*, 2013, **1**, 7904–7913.
- 51 G. H. Roche, D. Flot, J. J. E. Moreau, O. J. Dautel, J.-S. Filhol and A. van der Lee, *Cryst. Growth Des.*, 2021, **21**, 3850–3863.
- 52 K. K. Jha, S. Dutta and P. Munshi, *Cryst. Growth Des.*, 2018, **18**, 1126–1135.
- 53 A. Kumar, J. Chauhan, K. D. Dubey, S. Sen and P. Munshi, *Mol. Pharm.*, 2022, **19**, 1008–1018.
- 54 S. Dutta, V. Vikas, A. Yadav, R. Boomishankar, A. Bala, V. Kumar, T. Chakraborty, S. Elizabeth and P. Munshi, *Chem. Commun.*, 2019, **55**, 9610–9613.
- 55 G. Gupta, Y. Gupta, A. Kumar, R. Bhowal and P. Munshi, *Cryst. Growth Des.*, 2024, **24**, 646–656.
- 56 W. C. McCrone, *Chapter*, 1965, **8**, 725–767.
- 57 K. K. Jha, S. Dutta, V. Kumar and P. Munshi, *CrystEngComm*, 2016, **18**, 8497–8505.
- 58 D. Gentili, M. Gazzano, M. Melucci, D. Jones and M. Cavallini, *Chem. Soc. Rev.*, 2019, **48**, 2502–2517.
- 59 H. Chung and Y. Diao, *J. Mater. Chem. C*, 2016, **4**, 3915–3933.
- 60 R. Bu, H. Li and C. Zhang, *Cryst. Growth Des.*, 2020, **20**, 3561–





- 3576.
- 61 S. Dutta, Vikas, V. Thangavel and P. Munshi, *ACS Appl. Electron. Mater.*, 2021, **3**, 3633–3640.
- 62 G. B. A.M Glazer, *Space Groups for Solid State Scientists*, Elsevier, 2013.
- 63 V. Wadhawan, *Introduction to Ferroic Materials*, CRC Press, 2000.
- 64 H. Schmid, *J. Phys. Condens. Matter*, 2008, **20**, 434201.
- 65 H.-Y. Zhang, Y.-Y. Tang, P.-P. Shi and R.-G. Xiong, *Acc. Chem. Res.*, 2019, **52**, 1928–1938.
- 66 P. P. Shi, Y. Y. Tang, P. F. Li, W. Q. Liao, Z. X. Wang, Q. Ye and R. G. Xiong, *Chem. Soc. Rev.*, 2016, **45**, 3811–3827.
- 67 K. Aizu, *Phys. Rev. B*, 1970, **2**, 754–772.
- 68 S. C. Abrahams, E. Buehler, W. C. Hamilton and S. J. Laplaca, *J. Phys. Chem. Solids*, 1973, **34**, 521–532.
- 69 D.-W. Fu, H.-L. Cai, Y. Liu, Q. Ye, W. Zhang, Y. Zhang, X.-Y. Chen, G. Giovannetti, M. Capone, J. Li and R.-G. Xiong, *Science*, 2013, **339**, 425–428.
- 70 D.-W. Fu, W. Zhang, H.-L. Cai, J.-Z. Ge, Y. Zhang and R.-G. Xiong, *Adv. Mater.*, 2011, **23**, 5658–5662.
- 71 Z. S. Yao, K. Yamamoto, H. L. Cai, K. Takahashi and O. Sato, *J. Am. Chem. Soc.*, 2016, **138**, 12005–12008.
- 72 H. Y. Ye, Y. Zhang, D. W. Fu and R. G. Xiong, *Angew. Chemie - Int. Ed.*, 2014, **53**, 6724–6728.
- 73 S. Koval, J. Kohanoff, J. Lasave, G. Colizzi and R. L. Migoni, *Phys. Rev. B*, 2005, **71**, 184102.
- 74 T. F. Nova, A. S. Disa, M. Fechner and A. Cavalleri, *Science*, 2019, **364**, 1075–1079.
- 75 Y.-Y. Tang, J.-C. Liu, Y.-L. Zeng, H. Peng, X.-Q. Huang, M.-J. Yang and R.-G. Xiong, *J. Am. Chem. Soc.*, 2021, **143**, 13816–13823.
- 76 S. Horiuchi and Y. Tokura, *Nat. Mater.*, 2008, **7**, 357–366.
- 77 S. Horiuchi, Y. Tokunaga, G. Giovannetti, S. Picozzi, H. Itoh, R. Shimano, R. Kumai and Y. Tokura, *Nature*, 2010, **463**, 789–792.
- 78 S. Horiuchi, R. Kumai and Y. Tokura, *Adv. Mater.*, 2011, **23**, 2098–2103.
- 79 S. Horiuchi, F. Kagawa, K. Hatahara, K. Kobayashi, R. Kumai, Y. Murakami and Y. Tokura, *Nat. Commun.*, 2012, **3**, 1306–1308.
- 80 S. Horiuchi, K. Kobayashi, R. Kumai and S. Ishibashi, *Nat. Commun.*, 2017, **8**, 1–9.
- 81 A. S. Tayi, A. Kaeser, M. Matsumoto, T. Aida and S. I. Stupp, *Nat. Chem.*, 2015, **7**, 281–294.
- 82 S. Horiuchi, F. Ishii, R. Kumai, Y. Okimoto, H. Tachibana, N. Nagaosa and Y. Tokura, *Nat. Mater.*, 2005, **4**, 163–166.
- 83 S. Horiuchi, R. Kumai and Y. Tokura, *J. Mater. Chem.*, 2009, **19**, 4421.
- 84 K. Kobayashi, S. Horiuchi, S. Ishibashi, F. Kagawa, Y. Murakami and R. Kumai, *Chem. - A Eur. J.*, 2014, **20**, 17515–17522.
- 85 S. Horiuchi, R. Kumai and Y. Tokura, *J. Am. Chem. Soc.*, 2013, **135**, 4492–4500.
- 86 J. Harada, T. Shimojo, H. Oyamaguchi, H. Hasegawa, Y. Takahashi, K. Satomi, Y. Suzuki, J. Kawamata and T. Inabe, *Nat. Chem.*, 2016, **8**, 946–952.
- 87 J. Harada, N. Yoneyama, S. Yokokura, Y. Takahashi, A. Miura, N. Kitamura and T. Inabe, *J. Am. Chem. Soc.*, 2018, **140**, 346–354. DOI: 10.1039/D4MA00444B
- 88 H.-Y. Ye, Y.-Y. Tang, P.-F. Li, W.-Q. Liao, J.-X. Gao, X.-N. Hua, H. Cai, P.-P. Shi, Y.-M. You and R.-G. Xiong, *Science*, 2018, **361**, 151–155.
- 89 M. Szafranski, *J. Phys. Chem. B*, 2011, **115**, 8755–8762.
- 90 S. Seyedraoufi, E. D. Sødahl, C. H. Görbitz and K. Berland, *arXiv Prepr. arXiv2306.00363*.
- 91 Y. Y. Tang, P. F. Li, W. Q. Liao, P. P. Shi, Y. M. You and R. G. Xiong, *J. Am. Chem. Soc.*, 2018, **140**, 8051–8059.
- 92 T. Vijayakanth, A. K. Srivastava, F. Ram, P. Kulkarni, K. Shanmuganathan, B. Praveenkumar and R. Boomishankar, *Angew. Chemie - Int. Ed.*, 2018, **57**, 9054–9058.
- 93 R. A. Wisconsin, N. R. Goud, J. T. Damron and A. J. Matzger, *Angew. Chemie - Int. Ed.*, 2018, **57**, 9044–9047.
- 94 R. C. Haddon and A. A. Lamola, *Proc. Natl. Acad. Sci.*, 1985, **82**, 1874–1878.
- 95 D. Semmingsen and J. Feder, *Solid State Commun.*, 1974, **15**, 1369–1372.
- 96 G. A. Samara and D. Semmingsen, *J. Chem. Phys.*, 1979, **71**, 1401–1407.
- 97 Y. Moritomo, Y. Tokura, H. Takahashi and N. Mōri, *Phys. Rev. Lett.*, 1991, **67**, 2041–2044.
- 98 G. Gilli, F. Bellucci, V. Ferretti and V. Bertolasi, *J. Am. Chem. Soc.*, 1989, **111**, 1023–1028.
- 99 P. Gilli, V. Bertolasi, V. Ferretti and G. Gilli, *J. Am. Chem. Soc.*, 1994, **116**, 909–915.
- 100 G. Gilli and P. Gilli, *J. Mol. Struct.*, 2000, **552**, 1–15.
- 101 R. R. Choudhury and R. Chitra, *Cryst. Res. Technol.*, 2006, **41**, 1045–1048.
- 102 Z. Zikmund, P. Vaněk, M. Havránková, B. Březina, M. Člrmák and M. Vášša, *Ferroelectrics*, 1994, **158**, 223–228.
- 103 J. Kroupa, P. Vaněk, R. Krupková and Z. Zikmund, *Ferroelectrics*, 1997, **202**, 229–234.
- 104 D. Keil, H. Hartmann and C. Reichardt, *Liebigs Ann. der Chemie*, 1993, **1993**, 935–939.
- 105 M. Owczarek, K. A. Hujak, D. P. Ferris, A. Prokofjevs, I. Majer, P. Szklarz, H. Zhang, A. A. Sarjeant, C. L. Stern, R. Jakubas, S. Hong, V. P. Dravid and J. F. Stoddart, *Nat. Commun.*, 2016, **7**, 13108.
- 106 H. Peng, J.-C. Qi, X.-J. Song, R.-G. Xiong and W.-Q. Liao, *Chem. Sci.*, 2022, **13**, 4936–4943.
- 107 M.-M. Russew and S. Hecht, *Adv. Mater.*, 2010, **22**, 3348–3360.
- 108 W.-Q. Liao, Y.-L. Zeng, Y.-Y. Tang, H. Peng, J.-C. Liu and R.-G. Xiong, *J. Am. Chem. Soc.*, 2021, **143**, 21685–21693.
- 109 W. Liao, B. Deng, Z. Wang, T. Cheng, Y. Hu, S. Cheng and R. Xiong, *Adv. Sci.*, 2021, **8**, 2102614.
- 110 Z. Wang, C. Huang, J. Liu, Y. Zeng and R. Xiong, *Chem. - A Eur. J.*, 2021, **27**, 14831–14835.
- 111 Y.-Y. Tang, Y.-L. Zeng and R.-G. Xiong, *J. Am. Chem. Soc.*, 2022, **144**, 8633–8640.
- 112 Y. Du, C.-R. Huang, Z.-K. Xu, W. Hu, P.-F. Li, R.-G. Xiong and Z.-X. Wang, *JACS Au*, 2023, **3**, 1464–1471.
- 113 H. Zhang, H. Jiang, Y. Zhang, N. Zhang and R. Xiong, *Angew. Chemie Int. Ed.*, 2022, **61**, e202200135.
- 114 S. Horiuchi, R. Kumai and S. Ishibashi, *J. Mater. Chem. C*,





- 2021, **9**, 13739–13747.
- 115 P.-F. Li, W.-Q. Liao, Y.-Y. Tang, W. Qiao, D. Zhao, Y. Ai, Y.-F. Yao and R.-G. Xiong, *Proc. Natl. Acad. Sci.*, 2019, **116**, 5878–5885.
- 116 Y. Ai, Y.-L. Zeng, W.-H. He, X.-Q. Huang and Y.-Y. Tang, *J. Am. Chem. Soc.*, 2020, **142**, 13989–13995.
- 117 P.-F. Li, Y. Ai, Y.-L. Zeng, J.-C. Liu, Z.-K. Xu and Z.-X. Wang, *Chem. Sci.*, 2022, **13**, 657–664.
- 118 H.-Y. Liu, H.-Y. Zhang, X.-G. Chen and R.-G. Xiong, *J. Am. Chem. Soc.*, 2020, **142**, 15205–15218.
- 119 C. F. C. Fitié, W. S. C. Roelofs, M. Kemerink and R. P. Sijbesma, *J. Am. Chem. Soc.*, 2010, **132**, 6892–6893.
- 120 C. F. C. Fitié, W. S. C. Roelofs, P. C. M. M. Magusin, M. Wübbenhorst, M. Kemerink and R. P. Sijbesma, *J. Phys. Chem. B*, 2012, **116**, 3928–3937.
- 121 H. Anetai, Y. Wada, T. Takeda, N. Hoshino, S. Yamamoto, M. Mitsuishi, T. Takenobu and T. Akutagawa, *J. Phys. Chem. Lett.*, 2015, **6**, 1813–1818.
- 122 H. Anetai, T. Takeda, N. Hoshino, H. Kobayashi, N. Saito, M. Shigeno, M. Yamaguchi and T. Akutagawa, *J. Am. Chem. Soc.*, 2019, jacs.8b11222.
- 123 T. Akutagawa, H. Koshinaka, D. Sato, S. Takeda, S.-I. Noro, H. Takahashi, R. Kumai, Y. Tokura and T. Nakamura, *Nat. Mater.*, 2009, **8**, 342–347.
- 124 D. Miyajima, F. Araoka, H. Takezoe, J. Kim, K. Kato, M. Takata and T. Aida, *Science*, 2012, **336**, 209–213.
- 125 S. Furukawa, J. Wu, M. Koyama, K. Hayashi, N. Hoshino, T. Takeda, Y. Suzuki, J. Kawamata, M. Saito and T. Akutagawa, *Nat. Commun.*, 2021, **12**, 768.
- 126 D. Bordeaux, J. Bornarel, A. Capiomont, J. Lajzerowicz-Bonneteau, J. Lajzerowicz and J. F. Legrand, *Phys. Rev. Lett.*, 1973, **31**, 314–317.
- 127 C.-R. Huang, Y. Li, Y. Xie, Y. Du, H. Peng, Y.-L. Zeng, J.-C. Liu and R.-G. Xiong, *Angew. Chemie Int. Ed.*, 2021, **60**, 16668–16673.
- 128 S. K. Park and Y. Diao, *Chem. Soc. Rev.*, 2020, **49**, 8287–8314.
- 129 L. Li, P. Commins, M. B. Al-Handawi, D. P. Karothu, J. M. Halabi, S. Schramm, J. Weston, R. Rezugui and P. Naumov, *Chem. Sci.*, 2019, **10**, 7327–7332.
- 130 D. P. Karothu, R. Ferreira, G. Dushaq, E. Ahmed, L. Catalano, J. M. Halabi, Z. Alhaddad, I. Tahir, L. Li, S. Mohamed, M. Rasras and P. Naumov, *Nat. Commun.*, 2022, **13**, 2823.
- 131 A. Piecha-Bisiorek, A. Białońska, R. Jakubas, P. Zieliński, M. Wojciechowska and M. Gałązka, *Adv. Mater.*, 2015, **27**, 5023–5027.
- 132 N. Zhang, W. Sun, Y. Zhang, H.-H. Jiang, R.-G. Xiong, S. Dong and H.-Y. Zhang, *Nat. Commun.*, 2023, **14**, 5854.
- 133 H. Y. Zhang, Y. Y. Tang, Z. X. Gu, P. Wang, X. G. Chen, H. P. Lv, P. F. Li, Q. Jiang, N. Gu, S. Ren and R. G. Xiong, *Science*, 2024, **383**, 1492–1498.
- 134 <https://neutrons.ornl.gov/topaz>.
- 135 C. M. Fancher, C. M. Hoffmann, X. P. Wang, L. L. Daemen and A. J. Schultz, *APL Mater.*, 2021, **9**, 021111.
- 136 K. Kobayashi, S. Horiuchi, S. Ishibashi, Y. Murakami and R. Kumai, *J. Am. Chem. Soc.*, 2018, **140**, 3842–3845.
- 137 J. S. O. Evans, T. A. Mary and A. W. Sleight, *J. Solid State Chem.*, 1998, **137**, 148–160.
- 138 S. Margadonna, K. Prassides and A. N. Fitch, *J. Am. Chem. Soc.*, 2004, **126**, 15390–15391. DOI: 10.1039/D4MA00444B
- 139 D. J. Williams, D. E. Partin, F. J. Lincoln, J. Kouvetakis and M. O’Keeffe, *J. Solid State Chem.*, 1997, **134**, 164–169.
- 140 A. E. Phillips, A. L. Goodwin, G. J. Halder, P. D. Southon and C. J. Kepert, *Angew. Chemie*, 2008, **120**, 1418–1421.
- 141 A. E. Phillips, G. J. Halder, K. W. Chapman, A. L. Goodwin and C. J. Kepert, *J. Am. Chem. Soc.*, 2010, **132**, 10–11.
- 142 K. W. Chapman, P. J. Chupas and C. J. Kepert, *J. Am. Chem. Soc.*, 2006, **128**, 7009–7014.
- 143 S. G. Duyker, V. K. Peterson, G. J. Kearley, A. J. Ramirez-Cuesta and C. J. Kepert, *Angew. Chemie Int. Ed.*, 2013, **52**, 5266–5270.
- 144 D. J. Cram and J. M. Cram, *Science*, 1974, **183**, 803–809.
- 145 M. Zakrzewski, M. A. White and W. Abriel, *J. Phys. Chem.*, 1990, **94**, 2203–2206.
- 146 H. Birkedal and D. Schwarzenbach, *Angew. Chemie - Int. Ed.*, 2002, **41**, 754–756.
- 147 C. Serre, F. Millange, C. Thouvenot, M. Noguès, G. Marsolier, D. Louër and G. Férey, *J. Am. Chem. Soc.*, 2002, **124**, 13519–13526.
- 148 L. D. DeVries, P. M. Barron, E. P. Hurley, C. Hu and W. Choe, *J. Am. Chem. Soc.*, 2011, **133**, 14848–14851.
- 149 A. D. Fortes, E. Suard and K. S. Knight, *Science*, 2011, **331**, 742–746.
- 150 B. Dwivedi, A. Shrivastava, L. Negi and D. Das, *Cryst. Growth Des.*, 2019, **19**, 2519–2524.
- 151 L. Negi, S. Kumar, B. Dwivedi and D. Das, *Cryst. Growth Des.*, 2021, **21**, 1428–1433.
- 152 S. Chuskit, T. Sethi, P. Harsha and D. Das, *Cryst. Growth Des.*, 2023, **23**, 1336–1342.
- 153 H. T. Chifotides and K. R. Dunbar, *Acc. Chem. Res.*, 2013, **46**, 894–906.
- 154 S.-L. Li, T. Xiao, C. Lin and L. Wang, *Chem. Soc. Rev.*, 2012, **41**, 5950.
- 155 J. Borges and J. F. Mano, *Chem. Rev.*, 2014, **114**, 8883–8942.
- 156 D. Das, T. Jacobs and L. J. Barbour, *Nat. Mater.*, 2010, **9**, 36–39.
- 157 S. Bhattacharya and B. K. Saha, *CrystEngComm*, 2014, **16**, 2340.
- 158 M. K. Panda, T. Runčevski, S. Chandra Sahoo, A. A. Belik, N. K. Nath, R. E. Dinnebier and P. Naumov, *Nat. Commun.*, 2014, **5**, 4811.
- 159 A. Bousseksou, G. Molnár, L. Salmon and W. Nicolazzi, *Chem. Soc. Rev.*, 2011, **40**, 3313.
- 160 S. Brooker, *Chem. Soc. Rev.*, 2015, **44**, 2880–2892.
- 161 B. R. Mullaney, L. Goux-Capes, D. J. Price, G. Chastanet, J.-F. Létard and C. J. Kepert, *Nat. Commun.*, 2017, **8**, 1053.
- 162 D. Das, *Acta Crystallogr. Sect. B Struct. Sci. Cryst. Eng. Mater.*, 2021, **77**, 309–310.
- 163 J. F. Nye and R. B. Lindsay, *Phys. Today*, 1957, **10**, 26–26.
- 164 A. L. Kitaigorodsky, *Molecular crystals and molecules*, Elsevier, 2012, vol. 29.
- 165 J. D. Dunitz and A. Gavezzotti, *Chem. Soc. Rev.*, 2009, **38**, 2622.
- 166 L. Negi, A. Shrivastava and D. Das, *Chem. Commun.*, 2018, **54**, 10675–10678.



## Journal Name

## ARTICLE

- 167 S. Hasebe, Y. Hagiwara, T. Ueno, T. Asahi and H. Koshima, *Chem. Sci.*, 2024, **15**, 1088–1097.
- 168 S. Dutta and P. Munshi, *J. Phys. Chem. C*, 2020, **124**, 27413–27421.
- 169 C. Lind, *Materials (Basel)*, 2012, **5**, 1125–1154.
- 170 Y.-C. Yang, X. Liu, X.-B. Deng, L.-M. Wu and L. Chen, *JACS Au*, 2022, **2**, 2059–2067.
- 171 H. Takahashi and R. Tamura, *CrystEngComm*, 2015, **17**, 8888–8896.
- 172 Vikas, L. Negi and P. Munshi, *Cryst. Growth Des.*, 2024, **24**, 2533–2541.

View Article Online  
DOI: 10.1039/D4MA00444B



**Parthapratim Munshi, Ph.D., FRSC***Professor, Department of Chemistry*Email: [parthapratim.munshi@snu.edu.in](mailto:parthapratim.munshi@snu.edu.in)

Tel: +91 120 2667 130; +91 9810315940

**SCHOOL OF  
NATURAL  
SCIENCES**

23 June 2024

Sub: Data availability statements

No primary research results, software, or code have been included, and no new data were generated or analyzed as part of this review.

Sincerely yours

