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De novo drug design and biological evaluation of coumarin–pyrimidine co-drug derivatives as diabetic inhibitors: expanding a multi-algorithm approach with the integration of machine learning in pharmaceutical research†

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Optimization of hit compounds was carried out using the density functional theory. We introduce the coumarin-based co-drug by linkages (amide and oxime) of coumarins with pyrimidines to improve the pharmacokinetic activity of coumarin owing to the synergic effect of both collectively. Pharmacokinetic studies, including drug likeness, were performed using SWISSADMET, and toxicity was identified using a web-based server. Screened lead molecule **W₆** (**C**) was synthesized and characterized by NMR spectroscopy, and its anti-diabetic activity was evaluated by *in vitro* α -amylase inhibition. This study identified **W₆** as an active, new and more effective candidate for diabetes in drug discovery.

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1 Introduction

Diabetes mellitus is a metabolic disorder in which the human body is unable to maintain blood sugar levels. In modern times, the frequency of diabetes is increasing rapidly, and it has become an alarming situation for human health worldwide. Diabetes can cause prolonged complications owing to the deprivation of different blood vessels and organs. In the 21st century, diabetes mellitus is a major chronic disease. It is estimated that in 2015, 415 million people faced this disease, but the ratio has increased rapidly and will reach 642 million by 2040. According to the World Health Organization forecast, in 2030, diabetes will become the 7th leading cause of death worldwide.^{1–3}

The two primary types of diabetes are type I diabetes and type II diabetes. Type I diabetes mainly depends on insulin and is most commonly diagnosed in adults and children. In type I diabetes, the immune system destroys the insulin-producing beta cells in the pancreases and leads to the loss of the peptide hormone, which inhibits the transfer of blood glucose to different body tissues, such as muscle and liver. Type II diabetes is defined as an abnormality of the blood glucose level.² In hyperglycemia, blood glucose levels are high, which creates disturbances in the metabolism of protein and carbohydrates. A high level of glucose in the blood can damage the kidneys and eyes and cause different cardiovascular diseases.¹

Several anti-diabetic drugs are available in the market to target different anti-diabetic agents. It is necessary to target those anti-diabetic agents that show better pharmacological activity. In this study, we target four anti-diabetic agents (alpha amylase, aldose reductase, DPP-IV, and glycogen phosphorylase) in anti-diabetic drug design. The fundamental function of an alpha amylase inhibitor is to prohibit the consumption of carbohydrates in the small intestine, and it can be very helpful in curing diabetes.⁴ One of the dominant approaches is to inhibit DPP-IV for the treatment of diabetes. DPP-IV is a serine protease enzyme that is most commonly present in the kidneys and intestines. The function of DPP-IV is glucose metabolism, and it removes incretin from the body. The function of aldose reductase enzymes is to convert glucose to sorbitol through the polyol pathway. However, for the treatment of diabetes, the main approach is to inhibit

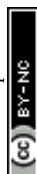
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these enzymes. The main function of phosphorylase enzymes is glycogen metabolism. In the case of type II diabetes, the amount of glucose is high in the liver. The main approach for the treatment of diabetes is to inhibit excess glucose production in the liver through glycogenesis.⁵

In ancient times, humans used natural products,⁶ such as plants, to treat different diseases. In recent years, natural products have played an important role in drug discovery. It is estimated that 25 percent of the drug was isolated primarily from the plants. Plants are a rich source of secondary metabolites.⁷ Antidiabetic drugs that are derived from natural sources, such as plants, have fewer side effects.⁸ It is a great challenge for scientists to identify drugs from natural sources that have fewer side effects.⁹

Coumarin, a natural product, is an attractive candidate for researchers owing to its fewer toxic effects, being more stable, high efficiency, and the fact that coumarin is a low molecular weight compound. Coumarin belongs to the class of benzopyrone, which consists of a benzene ring connected to the pyrone ring. The fundamental application of coumarin is its widespread use as a food additive, in cosmetics and the perfume industry, as well as in electrophoresis, hair spray, detergents, plastics, toothpaste, and food and beverages. Coumarin was first introduced to the market as an additive in 1868.¹⁰

The main purpose of this study is to introduce coumarin derivatives as a prodrug (mutual co-drug). A series of naturally occurring coumarin derivatives isolated from natural sources is identified. Coumarin-based co-drugs (mutual prodrugs) were introduced by linkages (amide, oxime) of coumarin derivatives with previously synthesized pyrimidine derivatives¹¹ by rational drug design. The advantage of a co-drug is that the absorption rate of the co-drug is better, and it can easily reach the target site.

Nowadays, *in silico* virtual screening methodologies play an important role in drug discovery. Molecular docking was performed to screen the compound. Molecular docking is a powerful tool for identifying the interaction of a drug with macromolecules, such as proteins. For virtual screening, molecular docking is the most widely used.¹² In computational studies, virtual screening plays a significant role. The fundamental purpose of virtual screening is to identify the chemical structures that show the best affinity for the targeted protein. Docking can be proposed to screen thousands of libraries in the compound. After screening the compound, optimization of the compound was done by DFT (density functional theory). Computational studies, such as DFT of screened compounds, give us information about the reactivity of the compound, geometry optimization and stability of the compound. The DFT of the screened compound was performed using Gaussian 09 software. The fundamental function of DFT is to determine the molecular structure of the compound and its activity by calculating different reactivity descriptors. By virtue of the information, the synthesis of the target can be identified as either possible or not. Pharmacokinetic studies, including drug likeness and toxicity, were performed using a web-based server. The ADMET study describes pharmacokinetic properties, such as

the absorption, distribution, metabolism, excretion, and toxicity of the compound.

2 Materials and methods

We explain here how the process unfolded in this study to provide a clear overview. At the outset, a set of known active coumarins and pyrimidines was assembled from the published literature. To assess whether these compounds interacted well with diabetic-related targets (α -amylase, DPP-IV, aldose reductase, and glycogen phosphorylase), they were used in molecular docking. These hits showing the highest performance were examined using Density Functional Theory (DFT) to check their electronic and chemical stability. Using SwissADME and Protox-II, the researchers carried out secondary checks on pharmacokinetics and toxicity to evaluate the drug-likeness of the synthetic compounds. **W₆**, the main compound, was made using several chemical procedures, identified with NMR and IR spectroscopy, and its ability to reduce diabetes was confirmed by testing with an α -amylase inhibition assay. Because various experimental and computational techniques were integrated, a favorable antidiabetic drug with minimal toxicity was discovered.

2.1. Software design for computational studies

Computational and virtual screening tools and software, such as ChemDraw Professional (Version 19.1.0.8, PerkinElmer, Waltham, WIA, USA), Gaussian 09W, Auto Dock Tools (ADT), BIOVIA Discovery Studio 2020 (v20.1.0.19295), the PyMOL molecules graphic system, version 2.4.1 (Schrodinger, LLC), and Open Babel 2.4.1, were utilized. Optimization of the compound was carried out using the basis set 6311G and B3LYP functions in Gaussian 09. A pharmacokinetic study was performed using a web-based server, such as Swiss ADMET (<https://www.swissadme.ch/>).

2.2. Rational drug designing

The application of rational drug design has gained the interest of medicinal chemists in drug discovery and medicinal chemistry owing to its ability to use theoretical and experimental knowledge of the underlying system. Rational drug design has contributed to the development of features and advancements in biochemistry, medicinal chemistry, molecular sciences, and pharmacokinetics of drugs. Based on rational drug design, knowledge of the system is utilized, which reduces the cost of human power, time consumption and experimental expenses in the lab.¹³ Therefore, based on the use of structure- and core-based designs, a series of coumarin and pyrimidine analogues was selected. Their inhibition potential was screened against type II diabetes using a multi-disciplinary approach of computational tools with the incorporation of *in vitro* antidiabetic analysis. For this purpose, a dataset of coumarin from the literature and pyrimidine previously synthesized¹⁴ were used to screen hit-to-lead optimized compounds. Lead compounds were further subjected to design using a core-hopping approach.



2.3. Preparation of the library

A library of 27 coumarin-based compounds was taken from the PubChem database, Google Scholar, DNP, and previously available literature. These coumarins were selected based on their *in vitro* activity, and pyrimidine derivatives used for co-drug were selected from the previously synthesized^{15,16} and *in vitro* and *in silico* optimized series.¹⁴ These selected compounds were collected in library form as a dataset (Table S1 (ESI)[†]).

2.4. Molecular docking analysis for diabetes inhibition of coumarin-pyrimidine co-drugs

2.4.1. Preparation of protein. Molecular docking was performed to determine the binding interaction of the ligand with the protein. Protein was prepared before docking the compound. For protein preparation, the water molecules were removed. From the protein data bank, the crystal structure was downloaded in PDB format (Table S2: ESI[†]). Preparation of the protein was done in the Discovery Studio. In the Discovery Studio, during protein preparation, all water molecules and heteroatoms were deleted. After selecting the ligand, the next step is to generate the binding sphere. Different values of the attributes of the binding sphere were noted. Protein was saved in PDB format. The protein was opened in Auto-Dock for docking purposes. In the AutoDock polar, a hydrogen atom was added. Different charges, such as Kollman charges and computed Gasteiger charges, were added. Protein was saved in pdbQT format in Auto dock.¹⁷

2.4.2. Ligand preparation. Ligand was downloaded from the PubChem database. The three-dimensional structure of the ligand was open in chem3D 19.1. In chem3D, different force fields were applied using a set of MM2 for energy minimization and MMFF94 for molecular dynamics. In chem3d, the three-dimensional structure was saved in a mol2 file. Using an Open-Babel mol2 file, the ligand was converted into a PDB format. The ligand was opened in AutoDock. A different step was performed, such as torsion tree; then, the ligand was saved in pdbQT format.

2.4.3. Grid preparation. The structure of the protein that targets different anti-diabetic agents was downloaded from the protein data bank. The binding site of all these selected proteins was determined by grid box preparation. The configuration of the grid box was set by determining the grid parameters *x*, *y*, and *z* according to different proteins. The grid box value for the alpha amylase inhibitor was *x* = 10.152200, *y* = 15.838836, and *z* = 41.117200. For the DPP-IV inhibitor, *x* = -9.767667, *y* = 8.120600, and *z* = 25.872900. For the aldose reductase inhibitor, *x* = -9.598192, *y* = 9.499346, and *z* = 17.858846, and for the glycogen phosphorylase inhibitor, *x* = 26.498438, *y* = 3.874062, and *z* = 51.496938. For all selected proteins, the size of the grid box was the same, such as (*x* = 15, *y* = 15, and *z* = 15); then, the grid box was saved in confi.txt. format.

2.5. MM-GBSA

The determination of the free binding energy of the protein-ligand complex is performed using the Molecular Mechanics-

Generalized Born Surface Area (MM-GBSA) method. This is done by employing the prime module of the Schrödinger suite 2022, which incorporates the OPLS4 force field and the VSGB (Variable Dielectric Surface Generalized Born) solvation model to perform the calculations.¹⁸ Following the energy terms and equation systems presented below, an automatic calculation was performed to determine the complex energies.

$$\Delta G_{\text{binding}} = G_{\text{complex}} - (G_{\text{receptor}} + G_{\text{ligand}}),$$

$$\Delta G_{\text{binding}} = \Delta E_{\text{MM}} + \Delta G_{\text{GB}} + \Delta G_{\text{SA}}.$$

Here, $\Delta G_{\text{binding}}$ is the receptor-complex binding free energy, while ΔE_{MM} is the gas phase energy in the force field of OPLS4. ΔG_{GB} and ΔG_{SA} are two electrostatic and polar free energy factors through the generalized Born (GB) and the non-polar solvation factor, respectively.¹⁹

$$\Delta G_{\text{binding}} = \Delta H_{\text{binding}} - T\Delta S,$$

where $T\Delta S$ characterizes the entropy drawback from binding and ΔH binding contains the enthalpy components, including van der Waals contacts, electrostatic bonds and solution effects. This study primarily estimates ΔG binding through enthalpy terms (ΔE_{MM} , ΔG_{solv} , and ΔG_{SA}) but eliminates the entropy component ($T\Delta S$) because normal mode analysis produces high computational demands. No direct evaluation of entropy contributions occurred because of computational constraints that matched common MM-GBSA protocol practices.

2.6. Molecular dynamics simulation

MD simulations evaluated the dynamic characteristics along with flexibility patterns in protein-ligand complexes. The iMODS server functions as an efficient tool for analyzing macromolecular complex intrinsic motions through Normal Mode Analysis (NMA). The key parameters included both deformability plots and *B*-factor analysis along with eigenvalues, covariance matrices and elastic network models. The computed eigenvalues describe how stiff different portions of the complex structures were, while deformability features together with *B*-factor analyses show where system flexibility occurs. The computed results validated both the structural stability and binding potential of the examined complexes when operating in dynamic environments.

iMODS is an online server used for the molecular dynamics' simulation of anti-diabetic protein-ligand complexes, such as 3RX2-W₆, 3BAJ-W₆, 1GPY-W₆, and 4J3J-W₆. NMA was carried out using the iMODS server to analyze the mobility and deformability of the complexes. Owing to this analysis, we discovered how the complexes move and flex as units, which is an important factor in understanding their ability to stay together and join others. Homologous molecules of the complexes could be processed by iMOD, but there is no need for 100% sequencing of the complexes. Simulation by iMOD plays a crucial role in drug designing strategies, as it gives knowledge about the flexibility of the target protein-ligand complex.



2.7. Density functional theory calculation

The computational studies used density functional theory and the Becke three-parameter Lee–Yang–Parr (B3LYP) functional with a 6-311G basis set to analyze the systems through the Gaussian 09W software package. The optimization of all molecules occurred in a gas phase scenario using an unrestricted optimization method. The structures reached true minima through frequency calculations, which showed no imaginary frequencies when performed at the same theoretical level. The optimized molecular structures yielded essential electronic values of HOMO and LUMO orbital energies and ΔE , μ , η , and ω that helped analyze the stability and reactivity features of these molecules.

The basic function of density functional theory is to calculate the structural parameters and orbital interactions of the chemical compound.²⁰ In recent years, density functional theory has played a significant role in determining the reactivity of organic compounds. DFT (density functional theory) calculation was carried out using a basis set 6311G and the B3LYP function. Optimization of the compound was carried out using density functional theory. Using Gaussian 09, we can determine the quantum chemical parameters, such as envision the dipole moment, electronic energy, molecular electrostatic potential and energies of HOMO and LUMO.²¹ The reactivity of the compound was determined by measuring the energy gap present between the HOMO and LUMO. Frontier molecular orbitals are the most important parameters for calculating the other chemical reactivity descriptors. However, the stability of the compound correlates with its hardness, and the chemical softness of the compound determines the range of its reactivity.²² However, the electrophilicity index indicates the flow of electrons between the donor and the acceptor. The compound has a higher value of electrophilic index. This indicates that the compound is more electrophilic in nature.²³ The compound has the highest value of the electrophilicity index, indicating that the compound has a higher binding affinity with organic molecules.²⁴ The compound with the higher value of chemical potential represents a moving charge from this compound to another compound.²⁵ The high dipole moment of the compound determines the binding site within the particular selected protein.²⁶ Molecular electrostatic potential is a fundamental tool for determining the reactivity and behavior of a compound. However, the MEP scale represents the electrophilic and nucleophilic regions. The red color indicates the electron rich region, and the blue color represents the electron deficient region. Molecular electrostatic potential indicates the size of the molecules. Molecular electrostatic potential represents the electron density in the compound.²⁷ The molecular electrostatic potential is determined using a basis set 6311G.²⁸

The synthetic reactants and the complex product formed were computationally optimized using Gaussian 09W and GaussView 6.0, with a hybrid functional of B3LYP and a basis set of 6311-G(d,p).²⁹ For parallel validation of the experimental details, calculations were carried out using the IEFPCM approach.³⁰

2.8. Drug-likeness activity of the selected compound

The drug-like property of the compound is determined by an ADMET study. Using ChemDraw, SMILES representations of the coumarin co-drug derivatives were generated and then processed on the SwissADME website to assess pharmacokinetic properties.³¹ The Lipinski rule of 5 was applied to determine the drug-like behavior of the compound. The toxicity of the compound was determined using the online server PROTOX-II.

2.9. Synthesis of target lead molecule

2.9.1. Synthesis of pyrimidine (A) as a precursor. A one-pot multi-component procedure (Scheme 1A) was adopted for the synthesis of the target compound with few modifications to the grindstone method. In this procedure, all reagents (in an equimolar ratio) were mixed in a round-bottom flask and stirred for 15 minutes. The reagents included ethyl acetoacetate, thio-urea, and 3-fluoro-benzaldehyde, along with a pinch of the $\text{CuCl}_2 \cdot \text{H}_2\text{O}$ catalyst. 3 ml HCl was added, and the mixture was stirred for 30 minutes to achieve ring formation. Then, the pyrimidine derivative (A) mixture was obtained and left overnight to achieve the maximum yield. Then, the mixture was washed with cold water to remove the acid and color of the catalyst, dried, washed, and recrystallized with polar solvents to obtain white crystals. TLC confirmed the formation of the product.

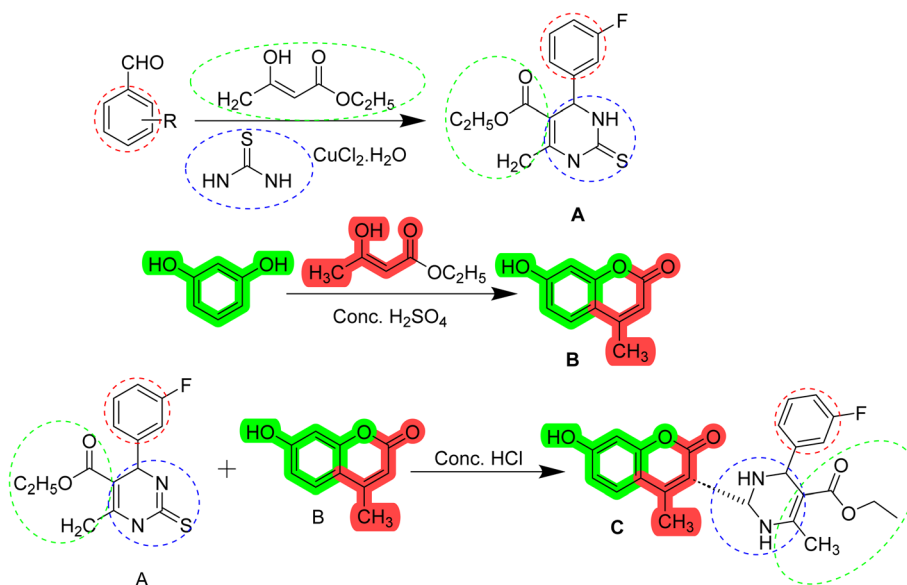
2.9.2. Synthesis of coumarin (B) for use as a precursor. For the synthesis of coumarin (B) (Scheme 1B), a solution of 5.5 g resorcinol dissolved in 10 ml ethanol and 6.5 g ethyl acetoacetate, and 6 ml H_2SO_4 was added dropwise at 0 °C. After the complete addition, the mixture was stirred for 5 hours at 40 °C. Then, 100 ml of cold water was added to the mixture and stirred for 2 hours. After 2 hours, precipitation was formed. The precipitates were washed, dried, and recrystallized to obtain white crystals.

2.9.3. Synthesis of lead co-drug from A and B. Both compounds A and B were mixed to achieve new compound C (Scheme 1C). This compound is a pyrimidine–coumarin complex. The solution of 4-methylumbelliferone (A) and pyrimidine derivative (B) in a round-bottom flask was stirred at 40–60 °C with the addition of 5 ml HCl after 15 minutes. After the HCl addition, the solution was stirred for 2 hours, with the formation of white precipitates. These precipitates were filtered, washed, and dried, and the ppt. was recrystallized to obtain the product crystals.

2.10. Spectroscopic characterization

The synthesized target lead was analyzed for structure confirmation by ^1H - and ^{13}C -NMR spectroscopy (see ESI† for spectral detail).

2.10.1. Characterization of A. White shiny crystalline solid; soluble in DMSO; mp = 202–206 °C; UV: λ_{max} 218, 325; FT-IR: 3163, 2811, 1893, 1879, 1852, 1802, 1672, 1614, 1451, 1391, 1370, 1336, 1328, 1312, 1275, 1240, 1227, 1216, 1160, 1135, 1069, 1018, 983, 898, 878, 865, 845, 806, 763, 755, 748, 694, 643, 584, 530 cm^{-1} ; ^1H -NMR (DMSO- d_6 , 300 MHz): & ^{13}C -NMR



Scheme 1 (A) Synthesis of the precursor pyrimidine (A) for C-drug formation; (B) synthesis of the precursor coumarin (B) for co-drug formation; and (C) synthesis of the target lead molecule (C) for co-drug formation.

(DMSO- d_6 , 25 MHz) (see ESI† for spectra), EIMS: m/z 294.02 (calcd for $C_{14}H_{15}FN_2O_5S$).¹⁴

2.10.2 Characterization of B. Colorless crystalline material; UV: $\lambda_{\max}$ 285.58; FT-IR: 332, 3010, 154, 1482, 1234, 1095, 550 cm^{-1} ; $^1\text{H-NMR}$ (DMSO- d_6 , 500 MHz) and $^{13}\text{C-NMR}$ (DMSO- d_6 , 100 MHz) (see ESI†); EIMS: m/z 176.02 (calcd for $C_{10}H_8O_3$).

2.10.3. Characterization of C. Colorless shiny crystalline material, soluble in DMSO; UV: $\lambda_{\max}$ 218, 284; FT-IR: 3327, 3163, 3010, 2811, 1893, 1879, 1852, 1802, 1672, 1614, 1451, 1391, 1370, 1336, 1328, 1312, 1275, 1240, 1227, 1216, 1160, 1135, 1069, 1018, 983, 898, 878, 865, 845, 806, 763, 755, 748, 694, 643, 584, 530, 154, 1482, 1234, 1095, 550 cm^{-1} ; $^1\text{H-NMR}$ (DMSO- d_6 , 500 MHz) & $^{13}\text{C-NMR}$ (DMSO- d_6 , 100 MHz) (see ESI†); EIMS: m/z 470.12 (calcd for $C_{24}H_{23}FN_2O_5S$).

2.11. Biological evaluation as *in vitro* anti-diabetic leads

The synthesized lead co-complex **W₆** (C) was subjected to bio-assay to screen its *in vitro* anti-diabetic potential using the method by Uzma Arshad *et al.*^{14,32} According to this method, the test compounds were incubated together with the α -amylase enzyme and starch under set conditions. Next, the amount of reducing sugars was determined by measuring their reaction with the 3,5-dinitrosalicylic acid (DNSA) reagent, most often at a wavelength of 540 nm, using a spectrophotometer. All patient outcomes were compared to those of the standard drug acarbose. To learn about their potential as anti-diabetic drugs, the amount of α -amylase inhibition by each test compound was measured against the control. The activity was performed using α -amylase as the enzyme and starch as the substrate in addition to acarbose as a reference drug. The percentage inhibition was calculated using the following formula:

$$A = \frac{A_{\text{control}} - A_{\text{sample}}}{A_{\text{control}}} \times 100.$$

3 Results and discussion

3.1. Scaffold-based drug design

Based on the core-hopping approach (Fig. 1), the selected series of compounds was screened using different computational tools. Consequently, one lead from pyrimidine and one from coumarin were further selected. These leads from both series were further subjected to synthesis, characterized for structure elucidation and further validated for their inhibitory potential by *in vitro* antidiabetic bio-assay.

3.2. Virtual screening of the compound based on the scoring function

Using a structure- and activity-based approach, coumarin-based compounds were downloaded from the PubChem database, which showed high IC_{50} values *via in vitro* and *in vivo* activities. A virtual screening of the compounds was performed to evaluate their biological potential as anti-diabetic agents. For this, a molecular docking approach was employed. Molecular docking of the compounds provided variable binding affinity values of the compounds. The screening of the compounds was achieved based on the scoring function of AutoDock software. Docking of the 27 naturally occurring coumarin compound-based pyrimidine hybrids was performed against target proteins, such as α -amylase with PDB-ID (3BAJ), DPP-IV with PDB-ID (4J3J), aldose reductase PDB-ID (3RX2), and glycogen phosphorylase PDB-ID (1GPY), which are related to diabetes. As hit compounds, they exhibited good binding affinity values when docked with AutoDock tools. Thus, a scoring function filter was applied, which provided the compounds with the best anti-diabetic potential with binding affinity values ranging from -10.1 to -9.0 kcal mol^{-1} .

Molecular docking studies of the compound with PDB ID 3BAJ showed a good binding affinity ranging from -9.7 to



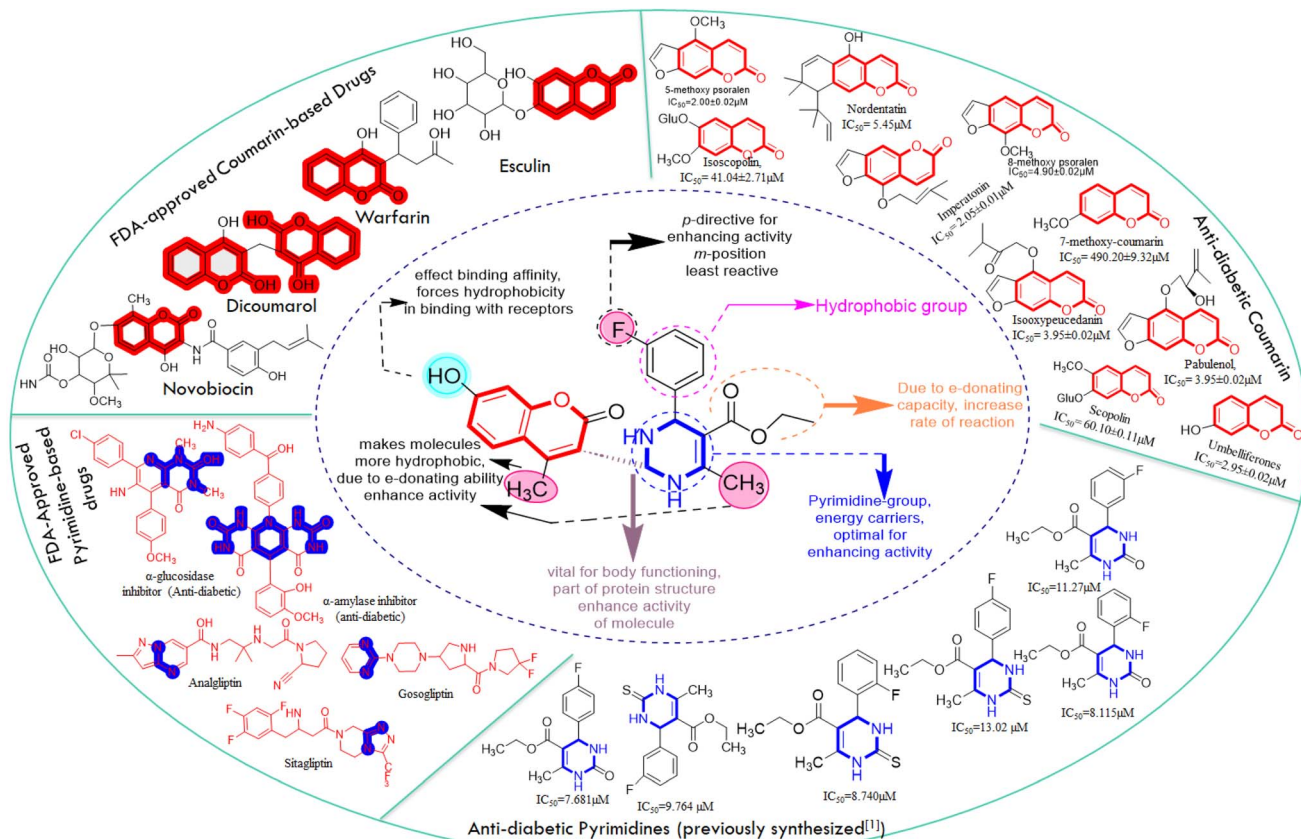


Fig. 1 Core-hopping and rational drug design-based approach used for the synthesis of C-drugs.

–7.0 kcal mol^{−1} compared to standard drugs. Acarbose showed a binding affinity of –6.9 kcal mol^{−1}. All the compounds showed a favorable interaction against the target enzyme.³³ The top five compounds showed a good binding score to the alpha amylase inhibitor shown in Table S3 (ESI).[†] Hydrogen bonding between the ligand and the targeted protein is depicted in Table S3 (ESI).[†] The compound that showed the highest binding score, 4-methyl umbelliferon–pyrimidine co-drug (**W**₆), showed hydrogen bonding interaction through amino acid residues, such as GLN63, HIS201, HIS305, HIS305, TRP59, and GLU233. It forms a six-hydrogen bond. Other Van der Waals interactions have been observed *via* pi-anion, pi-alkyl, alkyl, and pi-sulfur. In the 4-methyl umbelliferon–pyrimidine co-drug (**W**₆), electrostatic interaction was shown by ASP300, ASP197, and ASP300 amino acid residues, while hydrophobic interaction was shown by LYS200, ILE235, and HIS201 amino acid residues (Fig. 2).³⁴ None of the 27 compounds that target alpha amylase for anti-diabetic activity show any unfavorable interactions. Docking studies of 27 compounds with PDB-ID 4J3J for DPP-IV inhibition were performed, and only four compounds showed good binding affinity ranging from –9.5 to –9.1 kcal mol^{−1}. Sitagliptin, a standard drug, had a binding affinity of 8.7 kcal mol^{−1}. All the hit compounds formed 15 hydrogen bonds. More than 15 hydrophobic interactions were observed between the hit compound 8-methoxy psoralen–pyrimidine complex (**W**₂₄) and the amino acid residues LEU162, ILE235, HIS305, LEU162, LYS200, and ALA98. When the number of hydrogen bonds

decreased, it increased the hydrophobicity of the compound and indicated the stability of the compound (Fig. S1 and S2: ESI[†]).³⁵ Four hit compounds showed excellent binding affinity ranging from –10.1 to –9.0 kcal mol^{−1} compared to the standard drugs when docked with aldose reductase PDB ID 3RX2 (Table S1: ESI[†]). Epalrestat showed a binding affinity of –8.1 kcal mol^{−1}.³⁶ The hit compound (**W**₁) showed 12 hydrogen bonding interactions. The hit compound (**W**₁) showed hydrophobic interactions *via* amino acid residues Val47, leu300, Phe122, TRP219, Phe121, and others (Fig. 2).

Docking of 27 compounds showed binding affinity ranging from –9.5 to –8.8 kcal mol^{−1} when docked with glycogen phosphorylase protein 1GPY. The hit compound (**W**₁₉) showed higher binding affinity as well as four hydrogen bonds with binding protein 1GPY (Fig. S3: ESI[†]). However, metformin, which was used as a standard, showed a binding affinity of –5.7 kcal mol^{−1}.³⁷

3.3. MM-GBSA analysis

The performance accuracy of targeted proteins with selected lead ligands (selected from 27 studied complex compounds) was measured by the prime MM-GBSA calculation of receptor–ligand complexes.³⁸ The molecular mechanical energy produced by the receptor–ligand interaction was considered in the MM calculation. The effectiveness of this method has been demonstrated for a wide range of protein–protein



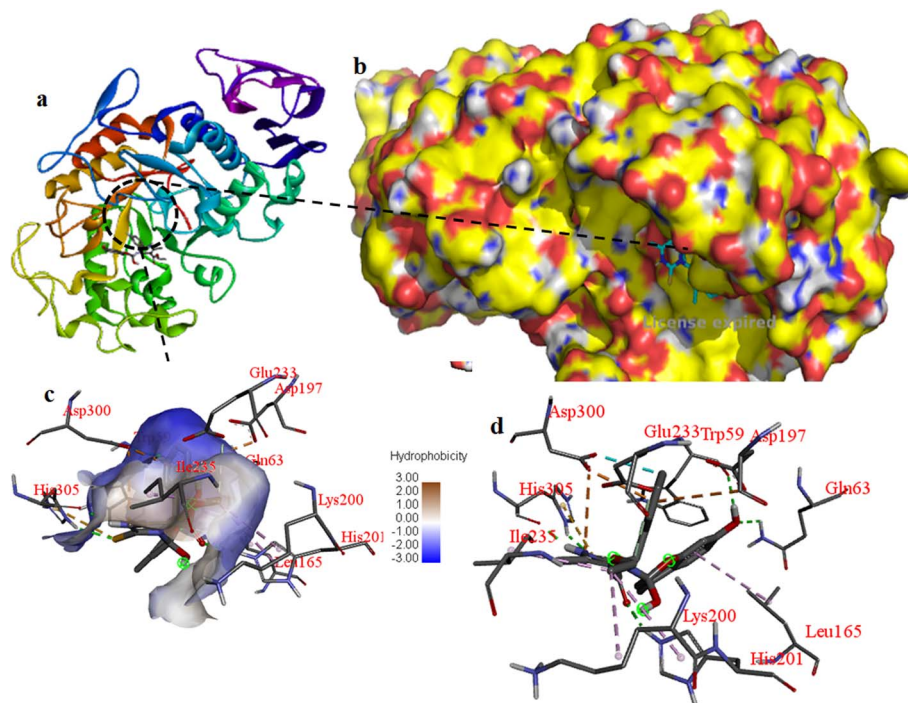


Fig. 2 Interaction of W6 with 3BAJ (a) protein in complex with ligand; (b) 3D-protein–ligand interaction, (c) 2D-protein–ligand interaction, and (d) hydrophobic interaction.

complexes.^{39,40} Molecular docking was also evaluated using MMGBSA free limiting energy, linked to the 1GPY, 3BAG, 3RX2 and 4J3J targets. Checking the expected low-energy poses is a surefire way to verify docking's precision. The Glide scores were quite close to the binding mode obtained experimentally using X-ray crystallography. Docking ligands into the binding pocket yields MM-GBSA free energy estimates.¹⁸ The free binding energies of the targeted complex with the respective

receptors are presented in Table 1. The linear relationship between binding free energy for different parameters of the complex in the solvent (water) is represented in Fig. 3.

3.4. Re-docking

An RMSD value was calculated to validate the docking method. Re-docking of the native and docked ligand was carried out for

Table 1 Calculated binding free energies of the targeted complex using prime module MM-GBSA, release of Schrodinger suite 2022

PDB ID	Ligand	ΔG_{bind}	ΔG_{bind} Coulomb	ΔG_{bind} covalent	ΔG_{bind} H-bond	ΔG_{bind} lipo	ΔG_{bind} vdW
1GPY	W ₁	−85.89	−19.92	0	−1.89	−28.67	−34.61
	W ₁₄	−63.85	−14.76	0	−3.51	−13.1	−46.4
	W ₁₉	−69.48	−41.7	0	−4.54	−13.97	−44.79
	Standard (metformin)	−37.65	−33.1	0	−3.34	−3.12	−24.12
3BAJ	Standard (acarbose)	−37.68	−23.28	17.03	−7.15	−27.5	−74.18
	W ₂₄	−46.54	−20	9.12	−2.29	−19.95	−63.37
	W ₆	−24.94	−16.12	−0.02	−1.96	−15.43	−41.42
	W ₃	−12.43	−9.28	−7.07	0.1	−18.26	−59.81
	W ₁₉	−28.43	−18.63	3.58	−0.5	−12.09	−55.54
	W ₂₃	−28.4	−17.09	4.6	0.02	−16.85	−50.1
	Standard (epalrestat)	−30.25	−6.41	14.83	−0.3	−11.1	−44.93
3RX2	W ₁	−39.06	−32.59	9.12	−3.11	−10.56	−32.44
	W ₃	−39.89	−1.57	4.8	−1.54	−17.31	−51.74
	W ₂₇	−58.65	−14.84	6.3	−0.94	−13.83	−58.29
	W ₂₃	−31.23	−6.96	6.27	−2.32	−14.14	−42.46
	Standard (sitagliptin)	−21.25	−13.75	9.75	−2.66	−20.81	−47.04
4J3J	W ₁	−33.57	−21.44	14.42	−2.37	−30.24	−46.2
	W ₃	−33.32	−9.83	11.54	−2.51	−26.51	−47.32
	W ₂₄	−42.89	−18.18	8.02	−1.13	−20.4	−64.13
	W ₁₈	−38.54	−17.61	16.57	−1.96	−18.39	−55.16
	W ₁₉	−42.46	−4.3	10.1	−2.4	−15.89	−39.87



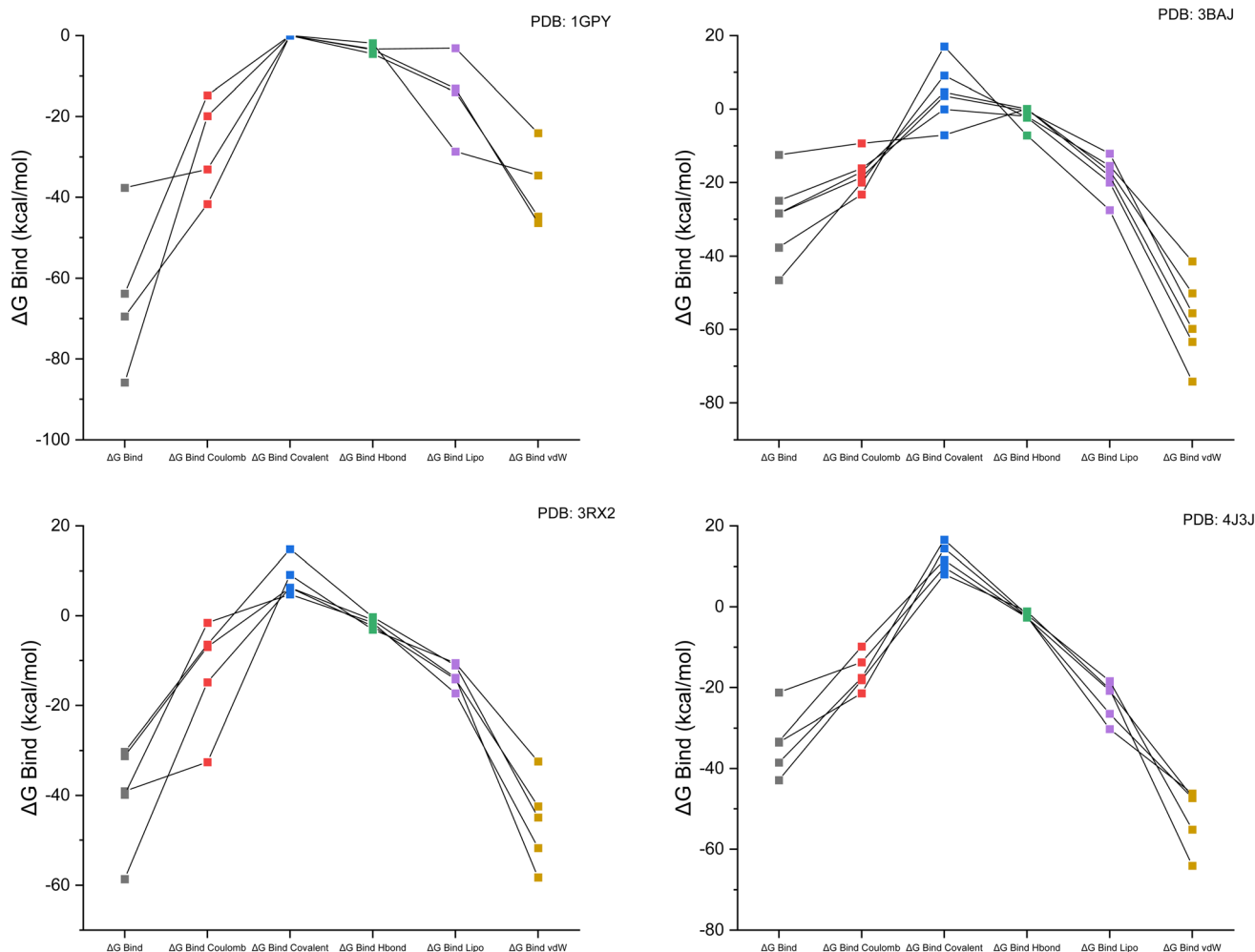


Fig. 3 Linear graphical representation of the binding free energy for Coulomb, covalent, H-bond, lipophilic and van der Waals forces for each receptor with the targeted ligand complex.

docking validation. Docking validation for antidiabetic activity was performed with PDB-ID 3BAJ for alpha amylase, 4J3J for DPP-IV, 3RX2 for aldose reductase, and 1GPY for glycogen phosphorylase, with an RMSD value less than 2 Å. The redocking best pose of the native and docked ligands is shown in Fig. 4.

3.5. Molecular dynamics simulation

All 4 complexes (3RX2- W_6 , 3BAJ- W_6 , 1GPY- W_6 , and 4J3J- W_6) undergo molecular dynamics simulation. The results from the iMODs summarize the NMA (normal mode analysis) for different antidiabetic docked complexes. This study was performed to regulate the actions of the proteins. The mobility of the docked complexes was checked using deformability and B -factor graphs. The deformability and B -factor graphs of complexes 3RX2- W_6 , 3BAJ- W_6 , 3RX2- W_6 , and 4J3J- W_6 are shown in Fig. S4 (ESI).[†] The eigenvalue of the complexes was linked to the normal mode of the complexes. The lower the eigenvalue, the higher the mobility of the complexes. The eigenvalues of all the complexes are presented in Table S4 (ESI).[†]

Complexes 1GPY- W_6 and 3BAJ- W_6 have lower eigenvalues, showing the greatest mobility of these complexes from the other two, as shown in Fig. S5 (ESI).[†] Complex 1GPY- W_6 has the lowest eigenvalue and great mobility compared to the other complexes. The RMSF (root mean square fluctuation) of lead complex 1GPY- W_6 was calculated by the iMODs, which showed the minimum fluctuation of 1.01×10^{-1} at residue no 128 and the maximum fluctuation of 9.87×10^{-2} at residue no. 718, as shown in Fig. 5.

According to the literature, the RMSD value for a good docked complex should be less than 2 Å, while the complex 1GPY- W_6 has an RMSD value of 0.25 Å, which is very good along with the radius of gyration that is also calculated by iMODs (Fig. 6).

3.6. DFT studies

Optimization of the screened compounds was carried out using DFT (Fig. 7). DFT was performed using a basis set of 6-311G.⁴¹ Frontier molecular orbitals can be classified as HOMO (highest occupied molecular orbitals) and LUMO (lowest unoccupied molecular orbitals) (Fig. 7). The frontier molecular orbitals are

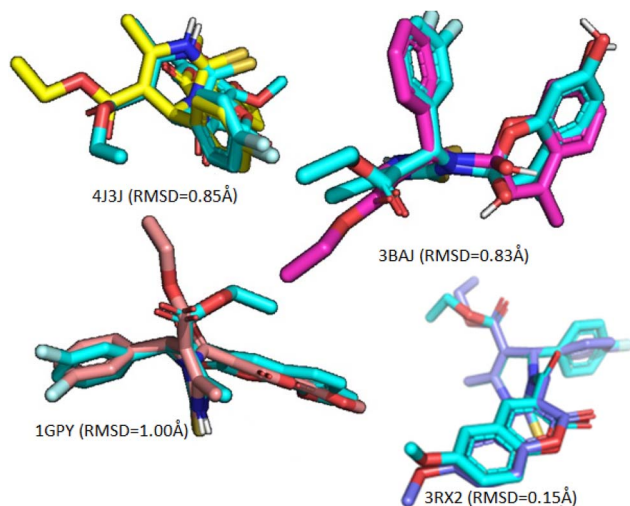


Fig. 4 Different redocking poses for the anti-diabetic activity.

very significant in determining the reactivity and stability of the compound.⁴² The compound with the highest HOMO value has the strongest ability to donate electrons. The compound has the lowest value of HOMO, and it has less ability to donate electrons, making it the least effective (Table 2).⁴³

Frontier molecular orbitals describe the electrophilic and nucleophilic characteristics of the compound. The compound with a higher energy of HOMO indicates that it is a good nucleophile, and the compound with a lower energy of LUMO indicates that it is a good electrophile.⁴⁴ Compound **W**₂₃ has the highest value of HOMO (−0.23429 eV), indicating that it is a good nucleophile and has a greater ability to donate electrons. The order of ranking of the nucleophilic character of the compounds is as follows:

$$\mathbf{W}_{23} > \mathbf{W}_{19} > \mathbf{W}_{18} > \mathbf{W}_{14} > \mathbf{W}_1 > \mathbf{W}_{24} > \mathbf{W}_{27} > \mathbf{W}_6 > \mathbf{W}_3.$$

Compound **W**₈ (−0.05721 eV) has a lower LUMO energy. It was found to have a good electrophile, meaning that it has a higher ability to accept electrons. The order of ranking of the electrophilic behavior of the compounds is as follows:

$$\mathbf{W}_3 > \mathbf{W}_1 > \mathbf{W}_6 > \mathbf{W}_{14} > \mathbf{W}_{27} > \mathbf{W}_{24} > \mathbf{W}_{23} > \mathbf{W}_{19} > \mathbf{W}_{18}.$$

The most important parameter is the energy gap, which is related to the stability of the compound. Compound **W**₇ has the lowest energy gap, which indicates the stability of the molecules. A small energy gap indicates that the compound is soft and more reactive. The ranking of the compound having the smallest energy gap and increasing the stability of the compound is given below:

$$\mathbf{W}_{27} > \mathbf{W}_{18} > \mathbf{W}_{19} > \mathbf{W}_{24} > \mathbf{W}_{23} > \mathbf{W}_{14} > \mathbf{W}_6 > \mathbf{W}_3 > \mathbf{W}_1.$$

Compound **W**₁ has a larger energy gap, which represents the hardness and less reactivity of the compound. The order of the decreasing reactivity of the compound is as follows:⁴⁵

$$\mathbf{W}_1 > \mathbf{W}_3 > \mathbf{W}_6 > \mathbf{W}_{14} > \mathbf{W}_{23} > \mathbf{W}_{24} > \mathbf{W}_{19} > \mathbf{W}_{18} > \mathbf{W}_{27}.$$

Chemical reactivity descriptors, such as chemical hardness and chemical softness values, quantify resistance to charge transfer. The electrophilicity index provides information about the binding affinity of compounds with organic molecules. The values of the chemical reactivity descriptor are shown in Table 3. The first component of the chemical reactivity descriptor is the electrophilicity index. Compound **W**₁ has the highest value of the electrophilicity index. This means that the compound has good binding affinity to the targeted protein. The order of compounds with good binding affinity is as follows:

$$\mathbf{W}_1 > \mathbf{W}_{18} > \mathbf{W}_{19} > \mathbf{W}_{23} > \mathbf{W}_{24} > \mathbf{W}_{27} > \mathbf{W}_{14} > \mathbf{W}_6 > \mathbf{W}_3.$$

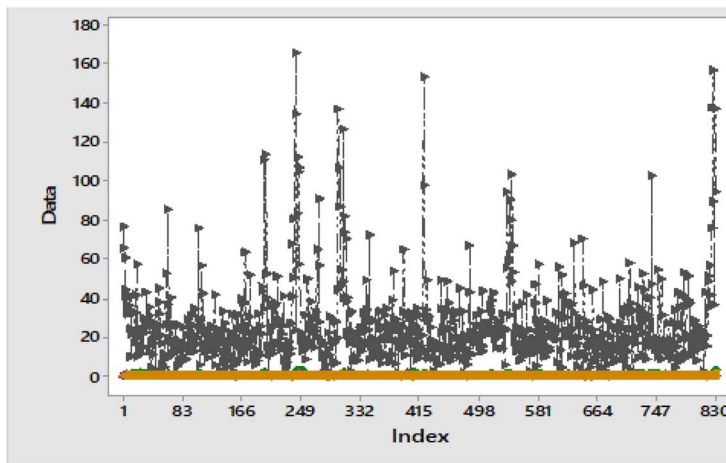
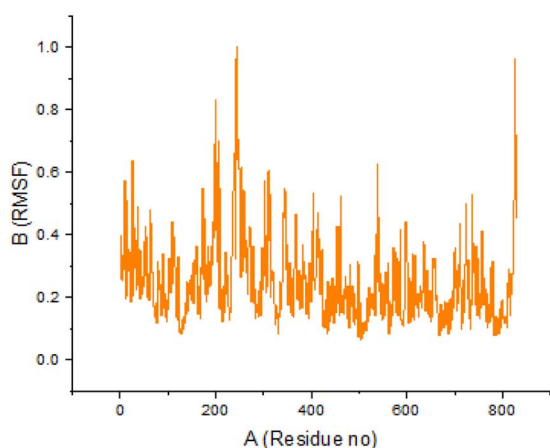


Fig. 5 RMSF and RMSF overlapping of the complex 1GPY-**W**₆.



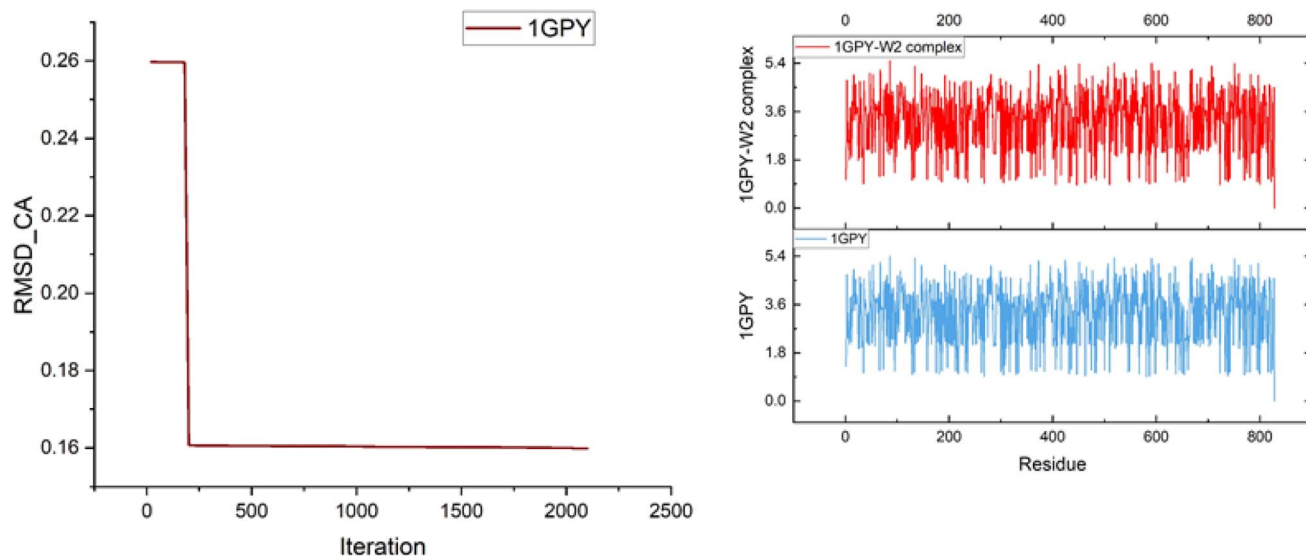


Fig. 6 RMSD and radius of gyration.

The chemical hardness and chemical softness of the compound are interrelated with the energy gap. The compound with a small energy gap indicates that the compound is reactive. However, hard molecules have a large energy gap and are the least reactive. Biologically, the activity of a compound increases when the compound has a small energy gap and the molecules are soft.

The compound has a small value of electronegativity, which determines the delocalization of electrons on the entire compound. This makes it easier for the compound to correlate with the biological system by giving electrons. The ranking of the compound based on the smaller value of electronegativity is as follows:

$$W_{27} > W_{18} > W_{19} > W_{24} > W_{23} > W_{14} > W_6 > W_3 > W_1$$

$$W_3 > W_1 > W_6 > W_{14} > W_{27} > W_{24} > W_{23} > W_{19} > W_{18}$$

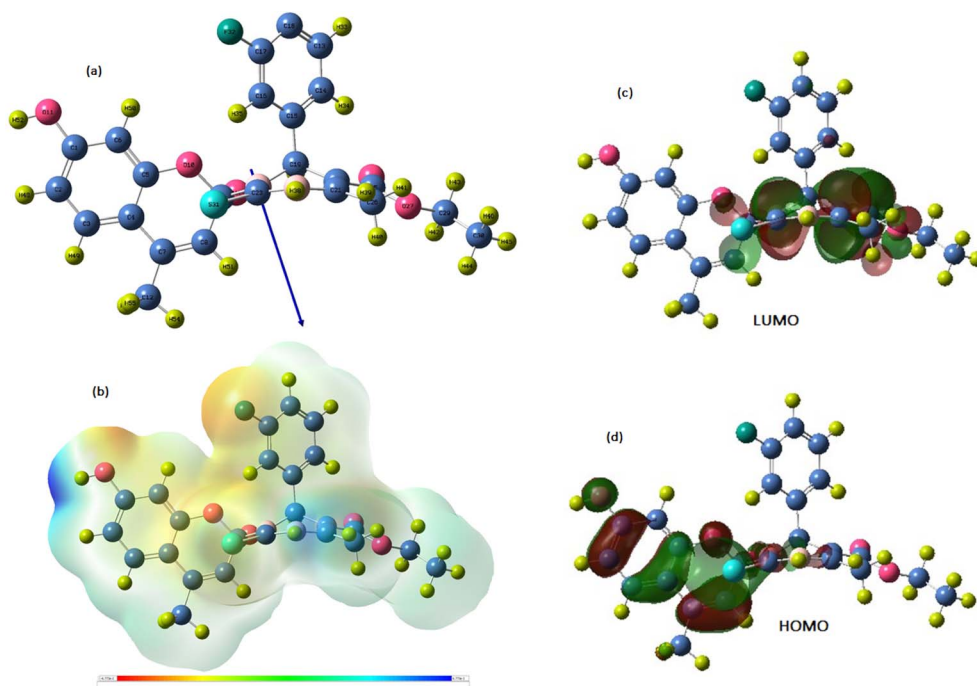


Fig. 7 (a) Optimized structure of the lead W_6 ; (b) MEP mapped of the lead W_6 ; (c) LUMO frontier molecular orbitals; and (d) HOMO frontier molecular orbital.

Table 2 Frontier molecular orbital values of hit compounds

Compound	E_{HOMO} (eV)	E_{LUMO} (eV)	Energy gap (ΔE)
W₁	−0.23	−0.06	0.17
W₆	−0.21	−0.07	0.14
W₁₄	−0.23	−0.09	0.14
W₂₃	−0.23	−0.10	0.13
W₁₈	−0.23	−0.11	0.12
W₁₉	−0.23	−0.11	0.13
W₂₇	−0.21	−0.10	0.12
W₃	−0.21	−0.06	0.15
W₂₄	−0.23	−0.10	0.13

The compound has the highest dipole moment, and it identifies the binding pose of the compound with the target protein. The order of the compound with the highest dipole moment is as follows:

$$\mathbf{W_{23} > W_{14} > W_{19} > W_{24} > W_{27} > W_6 > W_3 > W_{18} > W_1.}$$

Through DFT evaluations, we obtained essential data about the electronic structure and reactivity patterns in their target substances before conducting synthetic cycles. The compounds maintained effective HOMO–LUMO energy gaps that suggested optimal reactive qualities while ensuring chemical stability. Higher chemical hardness values showed resistance to deformation, thus predicting a favorable synthetic process. Electrophilicity displayed by the molecules confirmed their biological interaction ability while showing limited reactive potential, thus reducing nonspecific binding risks. Computational predictions built on quantum mechanics technologies allowed us to choose compounds for synthesis, which later served as the basis for molecular dynamics testing to establish the validity of our computational screening process.

3.6.1. UV/visible analysis of lead co-drug with precursors. A compound's electronic transitions indicate the specific part involved in the reaction and the structural conjugation.⁴⁶ The TD-DFT approach was applied with the B3LYP functional and 6311-G(d,p) basis set in DMSO as a solvent to fully comprehend the theoretical structural properties and excitation energies of the target compounds.⁴⁷ To assess the electronic characteristics of the compounds under study, singlet spin state excitations

were considered (Table 4). The solvent DMSO was used in the IEPCM model from the Gaussian calculation setup for theoretical calculations and in the experimental study; the same solvent was used. In this study, for compound **A**, which was used as a precursor, the maximum absorbance wavelength was 310.37 nm and λ_{max} of the experimental study was 325 nm, which proved that both computational and experimental results were in good agreement. Another precursor of this study was compound **B**; its λ_{max} (303.12) was calculated from the theoretical study, validating the results of the experimental study. The main target (co-drug) compound of this study, **C**, electronic excitations and maximum absorbance are also calculated for this compound to confirm and correlate it with experimental results. The maximum absorbance wavelength from the theoretical study was at 384 nm, and λ_{max} of the actual results was at 284 nm, which was quite different from the theoretical value. Excitation energies and oscillator strengths along with their λ_{max} were also calculated for all compounds mentioned in the table. A molecule with a HOMO LUMO band energy gap larger than 4 eV can absorb UV light, but one with less than 3 eV is predicted not to be a strong UV absorber (Fig. 8a–c and S6a–c: ESI†).^{48,49}

3.6.2. Vibrational analysis of target compounds. This analysis is used to identify the functional groups present in the particular compound that is under study. Theoretical methods were used to obtain the vibrational spectrum analysis of every studied compound to validate the results and peaks of the experimental spectrum. From the Gaussian calculation setup in Gauss View 6.0 of optimized compounds, a DFT with a 6-311G basis set along the B3LYP functional module was selected. The peaks of the compounds are discussed (Fig. 9a–c).

3.6.2.1. Aromatic ring C–H vibrations. In aromatic systems, the C–H of the benzene ring shows absorption at 3100–3000 cm^{-1} , as reported in the literature. In the theoretical study, compound **C** shows the C–H stretching of the benzene ring at 3100 cm^{-1} , but there is no peak observed in the experimental study. Compound **A** observed C–H stretching at 3106.81 cm^{-1} , and experimentally, this stretching appeared at 3163 owing to the presence of the alcoholic functional group moiety.

3.6.2.2. Carbonyl group (C=O) vibrations of ester linkage. In the literature, the ester-linked carbonyl functional group C=O stretching vibration range is 1750–1700 cm^{-1} . The main target compound (**C**) showed C=O stretching at 1700 cm^{-1} in the

Table 3 Different chemical reactivity descriptor values of hit compounds

Compound	Electronic energy	Dipole moment	Chemical potent	Ionization potential	Chemical hardness	Chemical softness	Electrophilicity index	Electronegativity	Electron affinity
W₁	−1705.56	2.25	−0.53	0.23	0.47	1.07	0.94	0.53	0.06
W₆	−1912.017	5.29	−0.53	0.21	0.47	1.07	0.31	0.53	0.07
W₁₄	−1946.73	6.40	−0.55	0.23	0.45	1.10	0.33	0.55	0.09
W₂₃	−2024.13	11.66	−0.55	0.23	0.45	1.11	0.34	0.55	0.10
W₁₈	−8592175.55	3.55	−0.55	0.23	0.45	1.12	0.34	0.55	0.55
W₁₉	−2061.04	5.90	−0.56	0.23	0.45	1.12	0.34	0.55	0.11
W₂₇	−2393.91	5.49	−0.55	0.21	0.45	1.11	0.33	0.55	0.10
W₃	−1858.37	4.318	−0.53	0.21	0.47	1.06	0.30	0.53	0.06
W₂₄	−2175.54	5.88	−0.55	0.23	0.45	1.11	0.34	0.55	0.10



Table 4 Experimental and theoretical calculations of the UV-vis for compounds A, B and C

Targets	Experimental λ_{max} (nm)	Theoretical λ_{max} (nm)	Excitation energies (eV)	Oscillator strength (f)
Compound A	325	310.37	3.9947	0.4399
Compound B	285.58	303.12	4.0903	0.3937
Compound C	284	389.65	3.1819	0.0114

theoretical study, which indicates the C=O moiety of ester linkage, but in the experimental study, this stretching was observed at 1752 cm^{-1} . The theoretical value of the same vibration for compound A appeared at 1705 cm^{-1} , but in the experimental study, no peak was observed in this range. In the vibrational analysis spectrum of compound B, no peak was observed in the theoretical and experimental calculations, which indicated that no ester linkage functional group was present in the particular compound.

3.6.2.3. Vibrational analysis of C-F. To assess the interaction of aromatic rings with halogen atoms from a functional group perspective, the stretching vibration for F-C is quite important. The detected infrared for this stretching is in the range of 1100–1300 cm^{-1} . The C-F stretching for compound C is stretched at 1250 cm^{-1} theoretically and experimentally; it was observed in the range of 1135 cm^{-1} . The theoretical and actual values of compound A for C-F stretching vibration were observed at 1127 cm^{-1} and 1135 cm^{-1} , respectively. From the theoretical and experimental results, no peak appeared for compound B of the same stretching, indicating that no halogen group was present in the B precursor compound.

3.6.2.4. Vibrational analysis of the C-OH and alcoholic OH groups. The O-H and C-O stretching vibrations have characteristic bands in the range of 3650–3400 cm^{-1} and 1300–1000 cm^{-1} , respectively, as reported in the literature. For target compound C, the O-H stretching appeared at 3327 cm^{-1} in the experimental results, but theoretically, no peak was observed in this range. The intense peak of the C-O stretching for compound C was observed at 1248 cm^{-1} theoretically and in the actual study; this vibration stretching appeared at 1018 cm^{-1} . In compound A, no peak was observed theoretically and experimentally in this range, clearly indicating that the alcoholic group was not present in compound A. Compound B showed the peak for O-H and C-O stretching at 3808 cm^{-1} and 1026 cm^{-1} theoretically, respectively.

3.6.2.5. Vibration analysis of C-N and N-H stretching. To identify the nitrogen group in the compound, N-H stretching in the infrared spectrum is very important. The N-H and C-N stretching of aromatic compounds are observed in the range of 3500–3100 cm^{-1} and 1250–1450 cm^{-1} , respectively, in the literature. No peak was observed in the range of N-H stretching for compound C, clearly indicating that no hydrogen was attached to the nitrogen atom in compound C. The C-N vibration stretching for compound C is observed at 1420 cm^{-1} in the theoretical study, while in the experimental study, the peak of the same stretching is observed at 1400 cm^{-1} . Compound A showed N-H stretching at 3163 cm^{-1} in the experimental

analysis, but in the theoretical analysis, this stretching was observed at 3106.81 cm^{-1} . The C-N vibration peak for compound A was observed at 1216 cm^{-1} and 1445 cm^{-1} by experimental and theoretical analyses, respectively. Further, no peak appeared for compound B in the range of both stretching vibrations C-N and N-H, indicating that no nitrogen group was present in compound B.

3.6.3. NMR characterization of the target compound.

Nuclear magnetic resonance spectroscopy is employed to examine the chemical surroundings, investigate molecular dynamics and ascertain the molecular structures. This study's main objective is to evaluate the GIAO technique of NMR using experimentally obtained values of hit compounds. The ^{13}C and ^1H NMR were calculated using DFT/B3LYP, and a 6-311G basis set was used in the input file of the optimized structure of the tilted compounds. The DMSO solvent was employed in both the GIAO and actual methods. If no comparable reference compound was chosen, conformers were not statistically treated, and the solvent effect, particularly for polar solvents, was not present; there would be some discrepancies and differences between theoretically estimated and experimentally obtained values for nuclear shielding. The reference utilized for all the target compounds in the computational study is TMS B3LYP/6-311 + G(d,p).

The highest theoretical chemical shift value of ^{13}C -NMR of co-drug C is 202.83 ppm, and the peak of this carbon was observed at 174.9 ppm. The irregularity in both peaks is due to the dependence on the surrounding functional groups present. The theoretical and experimental chemical shift values are close to each other, validating the experimental peaks, but the peak of C-20 appeared at δ 77.49 ppm in the GIAO method, but in the actual study, the peak of this carbon appeared at δ 98.6 ppm because of the presence of the pyran ring. The chemical shift values of ^1H NMR were found to agree with the experimental values (Table 5 and Fig. 10).

In compound B, the highest chemical shift value is at a chemical shift of 184.11 ppm in the theoretical study, and while in the experimental calculation, this carbon shows a peak at 161.1 ppm. In GIAO, the methyl carbon exhibits a peak at 37.26 ppm; however, in the experimental investigation, this carbon demonstrated a peak at 18.0 ppm. In compound A, the highest peak appeared at 174.61 ppm in the experimental study, while in the GIAO calculation, 168.93 ppm was the highest nuclear chemical shift value. There are no anomalies or differences between the theoretical and experimental values of the nuclear shielding of ^{13}C and ^1H for the compounds B and A.



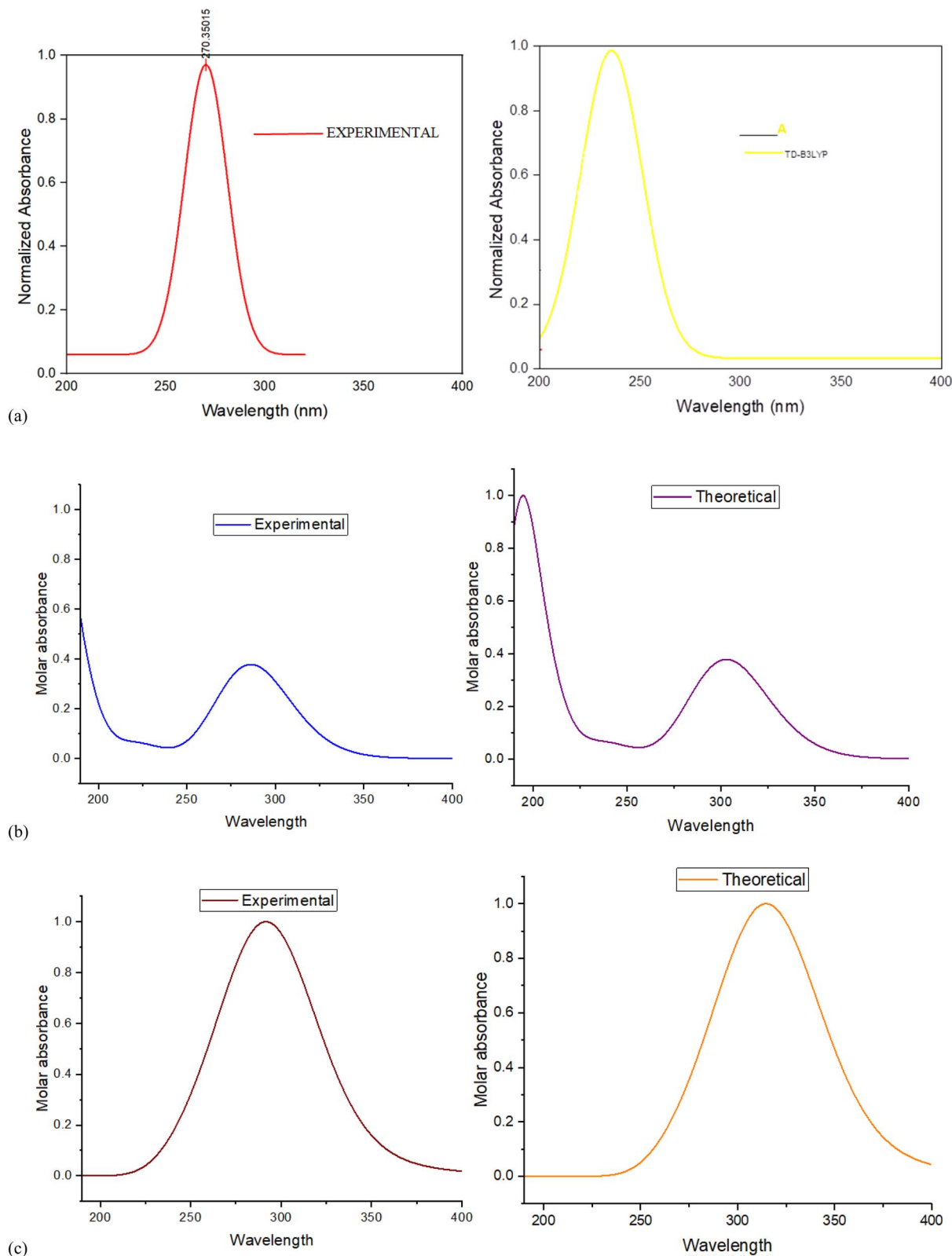


Fig. 8 (a) Plot of λ_{max} (nm) and molar absorbance of compound A. (b) Plot of λ_{max} (nm) and molar absorbance of compound B. (c) Plot of λ_{max} (nm) and molar absorbance of compound C.

3.7. ADMET studies

A pharmacokinetic study of the hit compound was conducted. The ADMET study consists of different parameters, such as

adsorption (*A*), distribution (*D*), metabolism (*M*), excretion (*E*) and toxicity (*T*). The toxicity of the lead compound was determined.



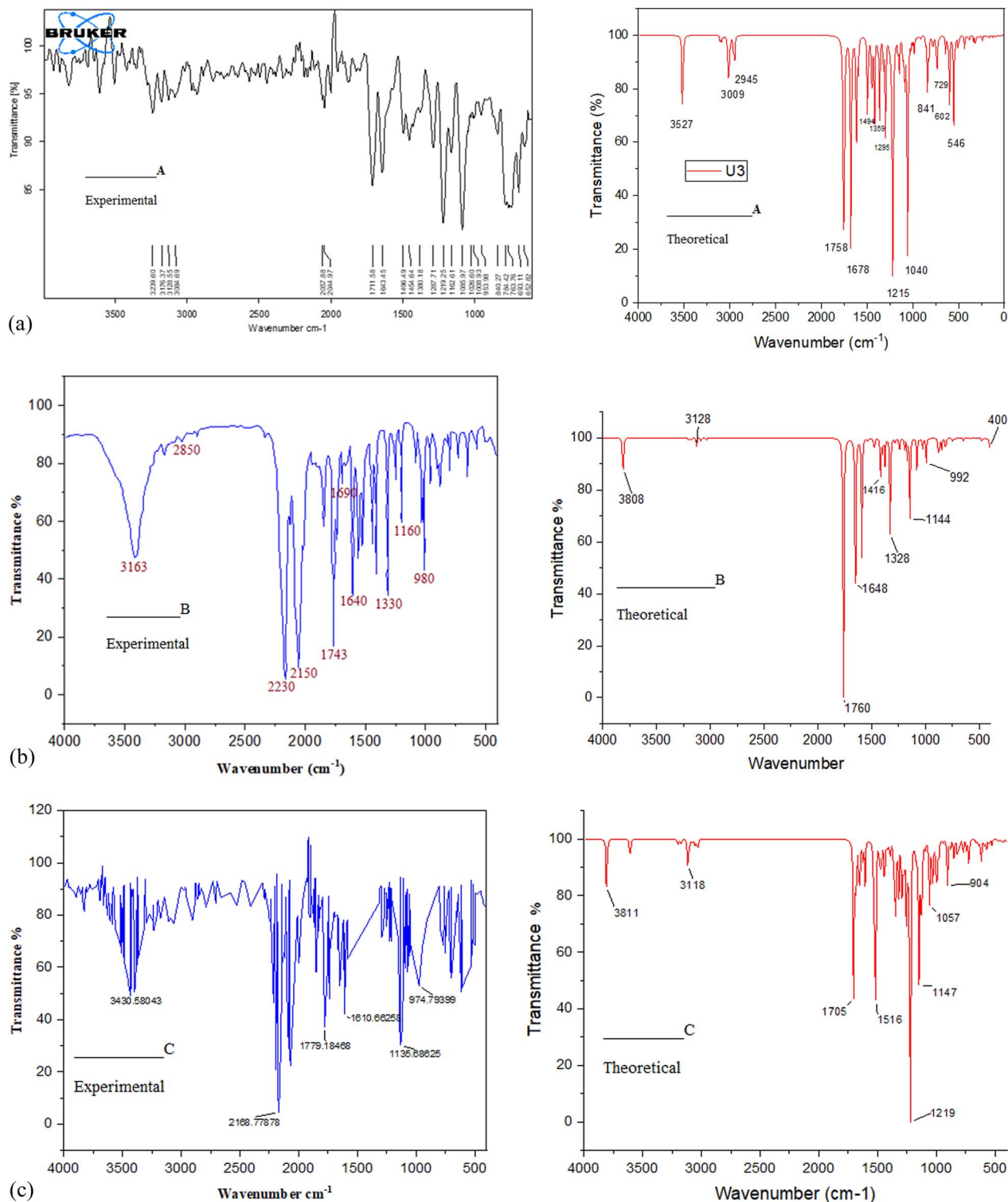


Fig. 9 (a) Experimental vibrational spectra and theoretical vibrational spectra of compound A. (b) Experimental vibrational spectra and theoretical vibrational spectra of compound B. (c) Experimental vibrational spectra and theoretical vibrational spectra of compound C.

The compound with a Lipinski rule of five with 0 violations indicated that the compound was orally active in the form of a drug in the human body. The compound with absorption percentage in the range of 84.9–100% indicates nontoxic.⁵⁰ The ADMET study revealed that W_6 , W_{23} and W_{14} fully obey the Lipinski rule with 0 (zero) violations. This indicates the drug-

like activity of the compound. However, the other compound showed 1 violation.

The molecular weight of the compound should be less than 500. The W_6 , W_{23} and W_{14} have low molecular weight compared to other compounds. However, the low molecular weight of the compound increases the absorption rate of the compound.⁵¹

Compounds **W**₆, **W**₂₃ and **W**₁₄ have small TPSA values. If the compound has a small TPSA value, it indicates that the compound can be easily absorbed. Except for one compound, all other compounds had a TPSA (total polar surface area) value less than 140. This indicates that these compounds passively pass through the cellular path.⁵² If the compound has more than 5 hydrogen bond donors and more than 10 hydrogen bond

acceptors, it does not obey the Lipinski rule. All the compounds have a number of hydrogen bond donors of less than 5, and a hydrogen bond acceptor of less than 10. The medicinal properties of these compounds indicate that they show good bioavailability. If the value of consensus log *P*_{o/w} is smaller than 5, it demonstrates that a suitable quantity of drug could be reached and shows aqueous solubility inside the body. All the

Table 5 Theoretical ¹H-NMR spectra and ¹³C-NMR spectra of compounds **A**, **B**, and **C**

Carbon	Theoretical	Experimental	Proton	Theoretical	Experimental
Compound A					
5-C	1.38	14.46	26-H	8.23	10.43
15-C	11.31	17.67	27-H	7.99	9.75
10-C	12.38	53.99	25-H	7.83	7.44
8-C	30.07	60.18	24-H	7.75	7.08
3-C	31.46	100.74	22-H	7.46	7.15
12-C	47.05	122.87	23-H	6.75	6.97
13-C	55.15	113.44	21-H	5.64	5.21
9-C	63.42	113.73	34-H	4.73	4.05
11-C	63.70	131.27	35-H	4.58	2.50
2-C	75.73	164.21	29-H	3.64	1.12
20-C	117.13	146.51	28-H	3.17	
1-C	121.13	160.97	30-H	2.13	
18-C	160.68	174.61	33-H	2.06	
19-C	168.93	165.43	32-H	2.04	
			31-H	1.83	
Compound B					
1-C	184.11	161.1	15-H	8.33	10.5
9-C	180.60	160.2	16-H	7.55	7.57
5-C	179.85	154.8	14-H	7.47	6.79
10-C	178.98	153.4	18-H	6.68	6.68
3-C	148.54	126.5	17-H	5.60	6.10
4-C	134.89	112.8	20-H	3.12	2.99
2-C	131.46	110.2			
11-C	130.41	112.0			
6-C	122.08	102.1			
13-C	37.26	18.0			
Compound C					
15-C	68.8	69.2	47-H	7.542587	9.24
28-C	188.78	163.2	48-H	7.51	7.78
23-C	187.15	165.2	45-H	7.25	7.39
3-C	182.55	147.6	43-H	7.23	6.99
1-C	176.83	147.6	36-H	7.23	9.24
21-C	169.01	163.2	46-H	7.16	7.09
18-C	167.65	163.2	34-H	6.48	5.15
7-C	160.07	160.8	44-H	6.46	7.36
25-C	153.03	160.8	35-H	6.45	7.05
5-C	146.98	113.0	37-H	4.73	6.99
26-C	145.29	130.5	38-H	4.54	4.03
22-C	138.08	122.2	55-H	4.10	4.00
8-C	137.54	98.6	51-H	2.38	2.24
24-C	136.35	114.1	41-H	2.05	2.25
6-C	134.35	130.4	39-H	1.86	1.09
4-C	127.76	98.6			
9-C	127.28	98.6			
19-C	126.70	122.1			
2-C	121.73	122.0			
33-C	82.90	98.6			
20-C	77.49	98.6			
31-C	37.49	40.1			
12-C	35.70	14.0			
32-C	31.10	17.7			



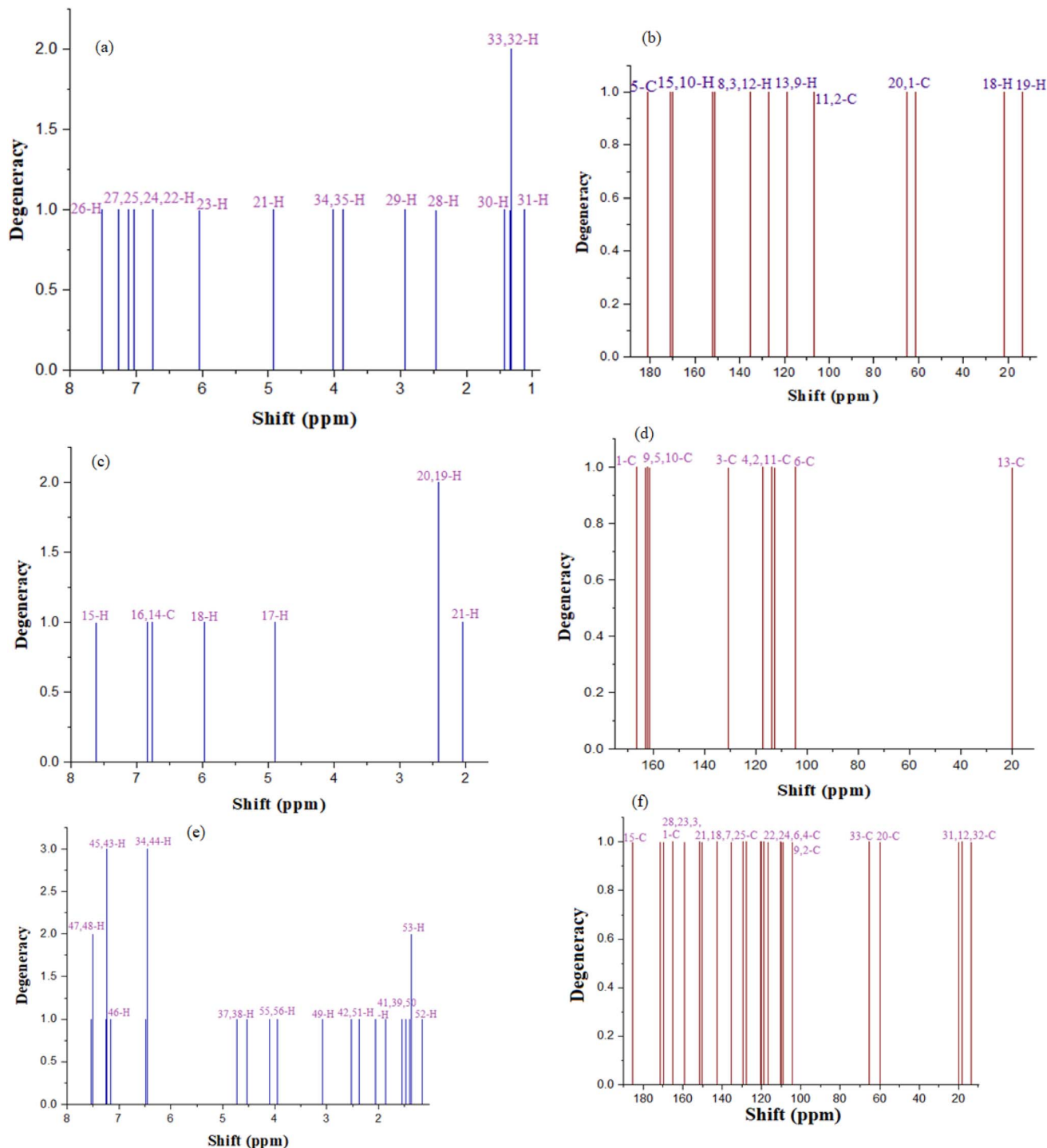


Fig. 10 (a) Theoretical ^1H -NMR spectra of compound A; (b) theoretical ^{13}C -NMR spectra of compound A; (c) theoretical ^1H -NMR spectra of compound B; (d) theoretical ^{13}C -NMR spectra of compound B; (e) theoretical ^1H -NMR spectra of compound C; and (f) theoretical ^{13}C -NMR spectra of compound C.

lead compounds have a value of less than 5.⁵³ The toxicity of the hit compound was determined using the online server. Out of the three compounds, only 1 compound showed the least toxic behavior owing to inactive carcinogenicity and immune-toxicity. The LD_{50} dose of these compounds is shown in Table 6, which is calculated from Protox II. The results indicate that a high dose of this compound was safe.

3.8. Chemistry of synthesized target molecules

The ^1H - and ^{13}C -NMR spectra of pyrimidine molecule **A** were found in accordance with the literature.¹⁴ Coumarin molecule **B** was elucidated by its ^1H - and ^{13}C -NMR spectrum analyses. ^1H -NMR of **B** displayed four protons in the aromatic region with a chemical shift value of 6.10 ppm as a singlet at position-3. A

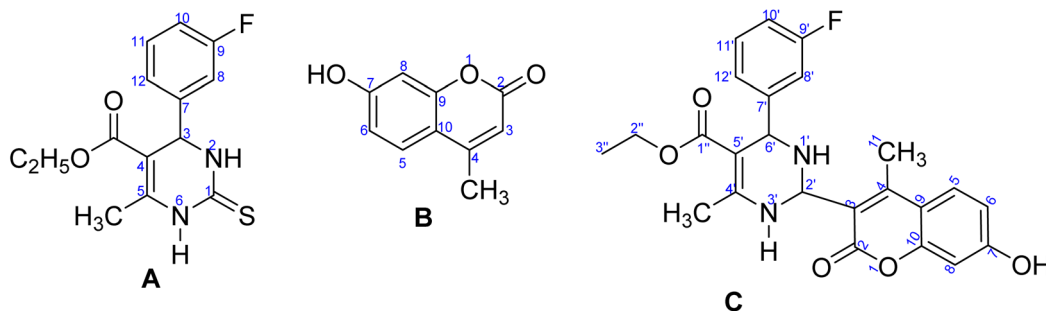
Table 6 Swiss ADME studies of hit compounds

Properties	Parameter	W ₆	W ₂₃	W ₂₄	W ₁₉	W ₁₄	W ₂₇	W ₁₈	W ₁	W ₃
Drug-likeness	Lipinski rule	0	0	1	1	0	1	1	1	1
	Ghose rule	No	No	No	No	Yes	No	No	No	No
	Veber rule	Yes	Yes	No	Yes	Yes	No	No	Yes	Yes
	Egan rule	Yes	Yes	No	No	No	No	No	Yes	Yes
	Mugge rule	Yes	Yes	Yes	Yes	Yes	No	Yes	Yes	Yes
Physico-chemical properties	Molecular weight	470.51 g mol ⁻¹	496.51 g mol ⁻¹	536.53 g mol ⁻¹	506.50 g mol ⁻¹	470.47 g mol ⁻¹	606.6 g mol ⁻¹	536.53 g mol ⁻¹	500.52 g mol ⁻¹	536.55 g mol ⁻¹
	No of H-bond donors	3	1	1	1	2	1	1	2	3
	No of H-bond acceptors	6	7	8	7	7	9	8	8	8
	TPSA	123.35	130.17	143.31	134.08	137.26	160.38	143	88.10	106.56
Lipophilicity of the compound	Consensus log P _{o/w}	3.24	3.53	3.82	3.63	2.73	4.40	3.86	4.06	3.80
	log P _{o/w} (XLOGP3)	3.22	3.87	4.27	4.30	3.17	5.07	4.27	4.24	3.95
	log P _{o/w} (WLOGP)	3.08	3.18	3.93	3.92	2.10	4.53	3.93	4.23	3.79
	log S (ESOL)	Moderate soluble	Moderate soluble	Moderate soluble	Moderate soluble	Moderate soluble	Poorly soluble	Moderate soluble	Moderate soluble	Moderate soluble
Solubility	log S (Ali)	-5.79	-6.30	-6.99	-6.83	-5.72	-8.18	-6.99	-5.80	-5.89
	GI absorption	High	High	Low	Low	High	Low	Low	High	High
	BBB permeant	No	No	No	No	No	No	No	No	No
	CYP1A2 inhibitor	No	Yes	Yes	Yes	No	Yes	Yes	No	No
Medicinal properties	Synthetic score	4.56	4.76	4.76	4.62	4.57	5.25	4.80	5.29	5.55
	Lead likeness	No	No	No	No	No	No	No	No	No
Toxicity	Immunotoxicity	Inactive	Active			Inactive				
	Cytotoxicity	Inactive	Inactive			Inactive				
	Mutagenicity	Inactive	Inactive			Inactive				
	Carcinogenicity	Inactive	Inactive			Inactive				
	LD ₅₀ mg kg ⁻¹	1000 mg kg ⁻¹	1190 mg kg ⁻¹			785 mg kg ⁻¹				

signal at δ 6.68 ppm was observed with a coupling constant of 2.5 and 6.0 Hz, which confirmed it at position-8 as *ortho*- and *meta*-coupled to another proton in close vicinity. For a doublet signal at δ 6.79 ppm with coupling constant 9.0, 2.5 Hz was observed at position-6, whose coupling confirmed it as *meta*- and *para*- to other protons. A proton at δ 7.57 ppm with a coupling constant of $J = 9.0$ Hz confirmed it at position-5. Among protons, a singlet at δ 10.5 ppm was found in the spectrum, which confirmed a hydroxyl at position-7, and

a methyl signal was found at δ 2.99 ppm as a singlet at position-4. The ¹³C-NMR data were found in accordance with ¹H-NMR data of molecule **B** with a distinguished peak at δ 160.2 for carbonyl, 161.1 for a carbon with a hydroxyl group and 18.0 for a methyl group. All methine and quarterly carbons were also confirmed by their respective signals in ¹³C-NMR spectra.

Lead molecule **C** (W₆) was obtained as a white crystalline solid. The ¹H- and ¹³C-NMR spectra showed characteristic peaks that help in identifying and confirming its structure. The ¹H-



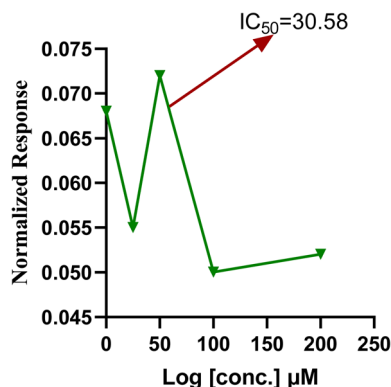


Fig. 11 Graphical representation of anti-diabetic activity and estimation of the IC_{50} value by non-regression analysis in GraphPad Prism.

NMR spectra of **W₆** displayed a characteristic peak at δ 9.24 ppm for a single amine, while the absence of another peak at δ 10.43 confirmed the formation of a junction of pyrimidine and coumarin. The presence of a characteristic peak at δ 130.4 and the absence of a peak at δ 160.2 for the carbonyl group also confirmed complex formation. These distinguished and characteristic peaks confirmed the formation of our target molecule. All other peaks of coumarin and pyrimidine rings in the spectra of **W₆** were found in accordance with the literature.

3.9. Antidiabetic evaluation of target molecule **W₆** (C)

Anti-diabetic results of the target molecule **W₆** (C) showed excellent inhibition potential with an IC_{50} value of 30.58 (Fig. 11). However, in comparison, acarbose was used as a standard and found to exhibit 61.40% inhibition. At various concentrations, **W₆** was screened for its potential against diabetes type II and was found to be more active with an IC_{50} value of 30.58 μM . *In vitro* antidiabetic activity was found in close correlation with binding affinity capacity calculated by molecular docking.

4 Conclusion

Coumarin and its analogues have different biological activities. Owing to their wide range of biological applications, coumarin and its analogues are of great interest to researchers because of their less toxic behavior. The main focus of this study was to identify targets as antidiabetic agents using coumarin derivatives and their analogues by interacting with pyrimidine compounds *via* amide and oxime linkages. A computational study was performed, and optimization of the molecular geometry of 27 compounds was performed by Density Functional Theory (DFT) using a basis set 6311G with the function B3LYP. Virtual screening of these compounds was carried out by molecular docking using the Auto-Dock tool. Structure-based virtual screening provided us with nine hit compounds with oxygen oxygen-centered furan rings and methoxy derivatives. These hit compounds showed good binding affinity for antidiabetic protein active sites. ADMET studies of hit compounds were performed, which filtered compounds that did not fall into

the criterion of drug likeness. Only 1 compound, **W₆** (C), fully obeys the Lipinski rule and is the least toxic, indicating drug-like behavior. Based on these screening results, lead molecule **W₆** was synthesized, characterized for structure and evaluated for antidiabetic potential by *in vitro* analysis. The *in silico* and *in vitro* results were found to be in close co-relation with each other, indicating that **W₆** (C) was a new effective candidate for diabetes inhibition.

Data availability

Data will be made available upon request.

Author contributions

Nusrat Shafiq: conceptualization, paper drafting, writing, supervision, molecular docking, NMR elucidation and funding; Warda Sarwar: library preparation, drafting, molecular docking, ADMET, and DFT; Maryam Rashid: supervision and editing; Nida NAz: synthesis of starting material and target molecule; Fatima Rida: MM-GBSA properties; Anika Moveed: DFT analysis; Dalal Hussien M. Alkhalifah: editing and validation; Mohamed Mohany: funding acquisition; Wael N. Hozzein: visualization; Shagufta Parveen: evaluation of results, data analysis, and literature; Eman Nawaz: editing, evaluation, and figures resolution.

Conflicts of interest

There is no conflict of interest.

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