Nanoscale



REVIEW

View Article Online



Cite this: Nanoscale, 2025, 17, 4906

NH₂-MIL-125(Ti) and its functional nanomaterials – a versatile platform in the photocatalytic arena

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Titanium (Ti)-based MOFs are promising materials known for their porosity, stability, diverse valence states, and a lower conduction band (CB) than Zr-MOFs. These features support stable ligand-to-metal charge transfer (LMCT) transitions under photoirradiation, enhancing photocatalytic performance. However, Ti-MOF structures remain a challenge owing to the highly volatile and hydrophilic nature of ionic Ti precursors. The discovery of MIL-125 marked a breakthrough in Ti-cluster coordination chemistry. Combining it with NH₂ chromophores to form NH₂-MIL-125 enhanced its structural design and extended its activity into the visible light region. This review delves into the high-performance photocatalytic properties of NH₂-MIL-125, focusing on its applications in H₂O₂ and H₂ production, CO₂ and N₂ reduction, drug and dye degradation, photocatalytic sensors, and organic transformation reactions. The discussion considers the influence of the Ti precursor, coordination environment, synthesis process, and charge transfer mechanisms. Numerous strategic methods have been discussed to improve the performance of NH₂-MIL-125 by incorporating linker modification, metal node modification, encapsulation of active species, and post-modification for enhancing light absorption ability, promoting charge separation, and improving photocatalytic efficiency. Moreover, future perspectives include methods to investigate how the efficiency of NH2-MIL-125-based materials can be planned in promoting research by highlighting their versatility and potential impacts in the area of photocatalysis.

Received 14th September 2024, Accepted 2nd January 2025 DOI: 10.1039/d4nr03774j

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1. Introduction

The persistent challenges of energy exhaustion and environmental deterioration resulting from excessive fossil fuel use have long troubled humanity, prompting researchers to explore renewable, clean energy sources. Harnessing sustainable solar energy to eliminate pollutants and generate power is widely acknowledged as an efficient and cost-effective strategy for environmental purification and addressing the energy crisis.2 Mimicking the natural photosynthetic processes of green plants and microorganisms, photocatalysis has emerged as an eco-friendly and energy-efficient method to tackle energy and environmental challenges.3 After the groundbreaking work of Fujishima and Honda, who utilized TiO2, a photocatalyst for solar energy conversion, the field has seen significant advancements.4 Numerous studies have concentrated on inorganic semiconductors, including metal oxides,5 metal sulfides, graphene, quantum dots, LDH, perovskites tec. Recently, the novel semiconductor graphitic carbon has

Centre for Nano Science and Nano Technology, S 'O' A (Deemed to be University), Bhubaneswar—751 030, Odisha, India. E-mail: kulamaniparida@soauniversity.ac.in, paridakulamani@yahoo.com attracted substantial research interest.¹¹ However, conventional photocatalysts typically suffer from a limited light absorption range, besides also having insufficient active sites and an increased rate of exciton pair recombination, which result in subpar performance. Therefore, there is an urgent need to develop well-grounded, high-performing catalysts that are receptive to light.

Metal-organic frameworks (MOFs), constructed from flexible organic ligands, metal ions, or well-elucidated secondary building units (SBUs), represent an innovative family of porous materials unveiled in the last decade. 12,13 MOFs stand out for their exceptional porosity, expansive surface area, well-ordered crystalline structure, and remarkable customizability, which allows the exploitation of inherent characteristics of organic ligands and metal clusters to produce a wide range of multifunctional MOFs. 14-16 The shape and extent of these organic ligands significantly influence the structure and symmetry of MOFs, while metal clusters impact their chemical stability and functional attributes, enabling the development of multifunctional materials.17 Research on photoluminescence has indicated that metal clusters within MOFs function as quantum entities of inorganic semiconductors. Organic linkers act as sensing organs that activate the quantum entities through a linker to the metal charge transfer (LCCT) upon exposure to light. 18-21

Combined with their structural analogy to zeolites, MOFs are promising for photocatalysis, compared with conventional semiconductors. In a MOF structure, the highest occupied molecular orbital (HOMO) of the organic linker can be easily tuned to optimize solar energy absorption. When excited, the electron can be transferred to a specific site within the MOF, such as the metal center or the lowest unoccupied molecular orbital (LUMO) of the organic linker.²² The crystalline structure enables the optimization of the migration distance for photoexcited charges, with the travel distance and angles being governed by the arrangement of the constituent units.²³ MOFs offer significant advantages in photocatalysis: (i) the crystallinity of MOFs reduces the recombination of photogenerated electron holes (e⁻-h⁺), as structural defects often found in photocatalytic materials serve as recombination centers.^{24,25} (ii) The extensive porosity of MOFs offers supplementary active sites and channels for transporting catalytic substrate, enabling swift transfer and utilization of exciton pairs.^{26,27} (iii) MOFs' structural adaptability enables the incorporation of long-range wavelength-absorbing groups (e.g., -NH₂) that act as organic bridging ligands. This enhancement broadens the light response spectrum and augments the generation of e⁻/h⁺ pairs upon illumination.²⁸ (iv) The adaptable compound structure of MOFs allows their easy integration with compounds like photosensitizers and co-catalysts, forming hetero or Schottky junction architectures. These structures aid in generating and separating photogenerated eand h⁺.^{29,30} Consequently, MOF-based photocatalysts are considered highly promising candidates for real-world applications in photocatalytic technology.

Among the many MOFs, most are built from divalent transition metal ions, resulting in open frameworks with good porosity but often poor stability in aqueous media. Increasing the metal ion charge and polarizing ability strengthens the metal-ligand interaction, enhancing the chemical stability and integrity of the coordination polymer. 31,32 While trivalent metal ions yield porous, water-stable MOFs, tetravalent ionbased MOFs are rare due to the high reactivity of such ions, often leading to oxides or amorphous coordination polymers. Zr(w) is a notable exception, with UiO-66 (the first Zr-carboxylate MOF featuring a robust hexa-nuclear Zr-oxo-cluster) spurring significant advances in Zr(IV)-MOFs.33 Higher-valence metals differ from lower-valence ones in their larger charge-toionic-radius ratio, making them more reactive and prone to oxygen affinity, resulting in multi-nuclear SBUs with high coordination numbers and greater framework robustness.34,35 Titanium (Ti), situated immediately above zirconium in the periodic table, represents the ionic form of a typical tetravalent element. However, its significantly reduced ionic radius compared with Zr⁴⁺ gives Ti a stronger affinity for O₂. TiO₂ exhibits solubility only in immensely acidic conditions (pH < 0), and this stability is transferred to Ti-based MOFs constituting titanium oxo-carboxylate SBUs featuring robust Ti-O covalent bonds.³⁶ Therefore, titanium as a Ti⁴⁺ ion presents an appealing option for creating MOFs with excellent chemical stability, redox ability, and photocatalytic properties (Ti4+ can undergo photoreduction to Ti³⁺). These features have led to swift

advancement in Ti-based MOFs, with associated studies expanding significantly over the last decade.

Recently, Wei et al. developed a new Ti-MOF using an imine condensation-based covalent approach, which maintained the integrity of the titanium cluster structure, leading to enhanced crystallinity and thermal stability. This design reduced the bandgap and shifted the CB, boosting CO2 reduction efficiency to 46 µmol g⁻¹ h⁻¹, outperforming MOF-902.³⁷ Yao et al. synthesized LCU-505, a Ti-MOF incorporating a tetranuclear [Ti₂Tb₂] cluster and a tritopic ligand, combining titanium and terbium to enhance both structural stability and photocatalytic performance. The material exhibited a highly microporous framework with a BET surface area of 882.4 m² g⁻¹, providing large void spaces that are well-suited for gas adsorption.38 Li et al. introduced a chelating coordination modulation (CCM) approach for synthesizing Ti-MOF single crystals. By employing chelating ligands, they stabilized the Ti center and controlled the formation of large single crystals with a stable octahedral structure. This method enabled the growth of highquality Ti-MOFs with customized properties suitable for a range of applications. 39 Garcia-Baldovi et al. made a significant advancement in Ti-MOF photocatalysis with the development of MIP-177-LT, a titanium-oxo cluster-based Ti-MOF featuring formate-imprinted active sites. This design facilitated efficient hydrogen production from formic acid, achieving an impressive quantum yield of 22% in the UV region, significantly outperforming TiO₂ P25 in photocatalytic activity. 40 Additionally, a self-assembled TiO₂/Ti-BPDC-Pt heterojunction, constructed via surface pyrolytic reconstruction, demonstrated efficient charge separation and achieved high H2 evolution rates of 12.4 mmol g⁻¹ h⁻¹, greatly surpassing the performance of conventional TiO2- or MOF-based systems. 41 Zhao et al. synthesized a 2D Ti-MOF featuring a honeycomb-like structure in which Ni atoms were anchored to Ti(w)-oxo nodes. This arrangement created an efficient S-scheme charge transfer system when coupled with BiVO4, resulting in a 66-fold enhancement in CO₂ photoconversion compared with pure BiVO₄, highlighting its potential for high-efficiency CO₂ reduction. 42 In 2024, Ti-MOF@DATp, a core-shell heterojunction material designed for the simultaneous photocatalytic reduction of U(vi) and oxidation of tetracycline, demonstrated exceptional photocatalytic performance, achieving 96% U(vi) removal and 90% tetracycline degradation. This underscored the multifunctional capability of Ti-MOFs for environmental remediation approaches.43 Subsequently, the field has evolved gradually yet significantly, integrating new ideas and methodologies relative to other MOF-related research areas.

The identification of the exceptionally porous MIL-125 by Serre and co-workers, made from dicarboxylate linkers and Ti₈ clusters, sparked a surge in Ti-based MOF synthesis and expanded their photocatalytic applications. ⁴⁴ The well-ordered 3D architecture of MIL-125 includes octahedral and tetrahedral cages—with attainable pore sizes of 12.55 and 6.13 Å, respectively. The MOF exhibits increased photonic sensitivity owing to the nature of titanium-oxo clusters, highlighting the ability of Ti-based MOFs for photocatalytic applications. This

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titanium-based MOF stands out owing to exceptional adsorption, thermal stability, and the capacity to incorporate highly dense immobile Ti-sites within the porous framework.⁴⁵ However, a wide band gap ($E_{\rm g} \sim 3.72$ eV) restricted its functions to the Ultraviolet (UV) spectrum. 44 Incorporating N-atoms by introducing -NH2 groups to the organic linker narrows the photocatalyst's E_g and prevents crystal structure deterioration.46 Fu and co-workers applied this strategy to engineer MIL-125 for UV photoresponse by substituting 2-amino terephthalate (NH₂-BDC) for BDC, creating the visible-light active NH₂-MIL-125.⁴⁷ NH₂-MIL-125(Ti) is attractive owing to its noticeable visible-light response, availability, and nontoxicity.48 Composed of cyclic octamers of TiO2 and ATA linkers, it is isostructural with MIL-125(Ti). 49 The -NH2 group induces an LMCT band structure, enhancing light absorption, 50 with an auxiliary transition band up to 550 nm, extending its visible light activity.51 The superior electrondonating characteristics of aromatic functional groups lead to the splitting of the valence band (VB) into a higher-energy occupying state. In contrast, the conduction band (CB), consisting of 2p and 3d orbitals of O and Ti, respectively, remains unaffected. This alteration leads to a narrower bandgap, extending the photocatalyst's photoresponse into the visible light spectrum.⁵⁰ NH₂-MIL-125(Ti), one among the initial Tibased MOFs, was explored in photocatalysis, with applications in hydrogen generation,⁵² organic pollutant degradation,⁵³ and CO_2 reduction. 54-56 It exhibits an E_g spanning from 2.5 to 2.6 eV (theoretical), making it highly suitable for visible-light semiconductor applications.⁵⁷ However, similar to most single-component photocatalysts, NH2-MIL-125(Ti) faces challenges such as limited light response and high charge carrier recombination rates. Additionally, MOFs' poor stability and low electronic conductivity restrict their photocatalytic applications. 51 To surpass these constraints and enhance efficiency, the construction of heterojunction by integrating Ti-MOFs with other functional nanomaterials has gained traction.⁵⁰ This strategy can enhance the charge carrier's transfer rate, minimize the recombination, and boost overall light absorption intensity by pairing the MOF with narrow-band-gap semiconductor materials. 58-62 Furthermore, as a porous material, NH₂-MIL-125(Ti) can act as an adhesion carrier for other semiconductors, enhancing specific surface area and reducing composite agglomeration.⁶³ This porous structure facilitates charge conduction and reactive substance movement, such as pollutant or CO2 molecule mass transfer.

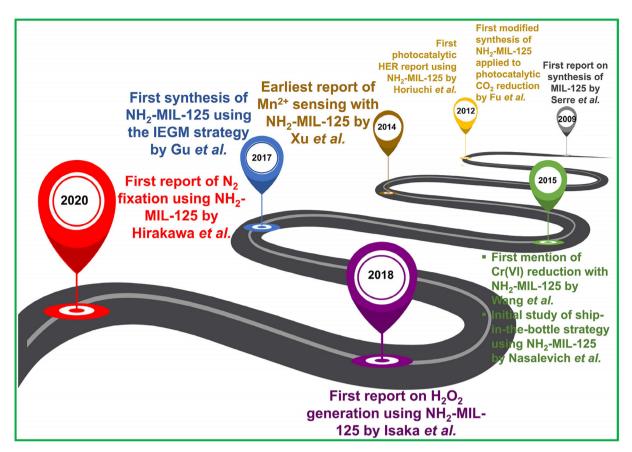
Over the past twenty years, there has been a proliferation of MIL-based photocatalysts, with significant emphasis on materials based on NH₂-MIL-125(Ti). These materials have been engineered for applications such as CO₂ reduction, water splitting, pollutant degradation, and organic transformation, leveraging their efficient absorption of visible light and robust stability. Numerous reviews have consolidated the applications of Ti-MOF-based photocatalysts. Assi *et al.* highlighted the key Ti-MOFs and their potential photoredox capabilities, while Yuan *et al.* comprehensively reviewed the structural arrangements, characteristics, and diverse applications of Ti-

MOFs. 68 These extensive works highlight the rapid development and significant potential of MOF-based photocatalysts. However, no review has focused entirely on NH2-MIL-125(Ti)based semiconductor photocatalysts. Due to undefined bottlenecks in photocatalysis and outdated or incomplete reports on NH₂-MIL-125(Ti)-based photocatalysts, there is an urgent need for a comprehensive and timely systematic review on this topic. In this review, we comprehensively outline recent significant advancements in NH₂-MIL-125(Ti) MOFs, particularly the role they play in the photocatalytic arena and related mechanisms. This review discusses unique topological structures, synthetic routes of NH₂-MIL-125(Ti) MOF, and various strategies to enhance photocatalytic efficiency. Moreover, it tackles significant challenges in ongoing research and outlines prospects for future investigations. Besides establishing a foundational repository for NH2-MIL-125(Ti), this comprehensive review also seeks to steer the blueprint and synthesis of Ti-based MOFs through an analysis of current frameworks. This methodology is poised to foster the exploration and advancement of more sophisticated MOF materials with possible applications within the discipline of photocatalysis. Scheme 1 illustrates the road map of the advancement of NH2-MIL-125(Ti) and its functional nanohybrid materials in the photocatalytic arena.

2. Chemistry behind titanium cations

Synthesizing Ti-based MOFs remains a highly challenging endeavor in the MOF research field, with reported Ti-based MOFs being relatively scarce compared with those involving other metals, which have only been recently discovered. This scarcity can be attributed to the unique chemistry behind Tications and the specific Ti-based complexes utilized as Ti precursors in MOF preparation. In this section, we aim to go deeper into the core issues surrounding the role played by titanium in MOF synthesis, present our perspective as to why titanium is less explored, and discuss how researchers are overcoming these challenges to produce more efficient Ti-MOFs.

The Ti(IV) ion is the most thermodynamically favored oxidation state, whereas Ti(III) can survive under reducing conditions, although it rapidly reverts to Ti(IV) when exposed to oxygen. However, Ti(IV) is susceptible to spontaneous hydrolysis in the presence of water, whether residual in solvents, intentionally added, or generated in situ during the reaction. This typically results in the formation of hefty multinuclear Tioxo clusters as secondary building units (SBUs), such as the octamer in MIL-125, along with the formation of amorphous TiO₂, even in highly acidic conditions. Due to its low electronegativity and high polarizability, TiO2 can partially block the pores of the formed MOFs. The robust stability of TiO2 makes it difficult to react with organic linkers, hindering the continuous formation of MOF architectures. This is illustrated in the titanium Pourbaix diagram (Fig. 1a), where Ti(IV) oxide predominates, and soluble molecular species are observed only in extremely acidic environments (pH < 0).67 Selecting the appropriate titanium precursor and complexing ligand, along with



Scheme 1 Timeline of the advancement of NH2-MIL-125(Ti) and its functional nanomaterials in photocatalysis to date.

the control of the pH of the media and the maintenance of a water-free condition, is critical for stabilizing Ti(IV) in solution and maintaining its reactivity. This shows that the greatest barrier in the field lies in maintaining a fine balance between stability and reactivity while working with titanium.

The titanium alkoxides, or Ti(OR)4 (where R is an alkyl group), are the most frequently employed Ti precursors for MOF preparation. Hydrophobic alkoxy groups and steric hindrance reduce Ti4+ hydrolysis, with reactivity adjustable by altering the alkoxy chain.66 Among these alkoxides, titanium isopropoxide (Ti(iOPr)4) is frequently used for Ti-MOF synthesis as opposed to TiCl₄, which is highly volatile and produces a white mist of TiO2 and HCl when exposed to humid air. Because of its vigorous hydrolysis and rapid release of HCl under optimal conditions, the use of TiCl₄ is uncommon.⁷¹ In contrast, Ti(iOPr)4, featuring hydrophobic alkoxide ligands attached to titanium centers, exhibits greater stability in air. Furthermore, Ti cations can form complexes with a variety of additional ligands (such as β-diketonates, carboxylates, phosphonates, catecholates, etc.), which, due to their bidentate nature, increase the coordination number and thereby enhance the stability of these species in aqueous environments. To prevent hydrolysis and improve the handling of the active titanium source, the entire synthesis process must be conducted in an inert glove box, presenting challenges for the

overall procedure. As a result, while this approach is effective for synthesizing Ti-MOFs, its complexity has limited its widespread adoption. Alternatively, pre-forming Ti-oxo clusters can decelerate the nucleation and crystal formation reaction, thereby preventing hydrolysis. These clusters can either maintain their original structure or undergo rearrangement through association or dissociation during MOF synthesis.⁷² The fact that Ti ions can exist in multiple oxidation states (Ti⁴⁺/Ti³⁺) and coordinate with a broad range of ligands only adds to the versatility of Ti in MOF chemistry.⁶⁷ From our perspective, the tendency of Ti-precursors to be unstable and extremely reactive represents an open question in the field. Substantial advances have been achieved in stabilizing Ti(IV) in solution by using conditions such as moisture-free conditions and inert atmospheres or by forming Ti-oxo clusters pre-synthesis to slow down hydrolysis kinetics. Much work, however, remains to be done to optimize these processes.

3. Coordination behavior of titanium ions

In designing and tailoring MOF structures and functions, understanding the coordination habits of metal ions-specifically how they coordinate with oxygen from carboxylates to Review (v)

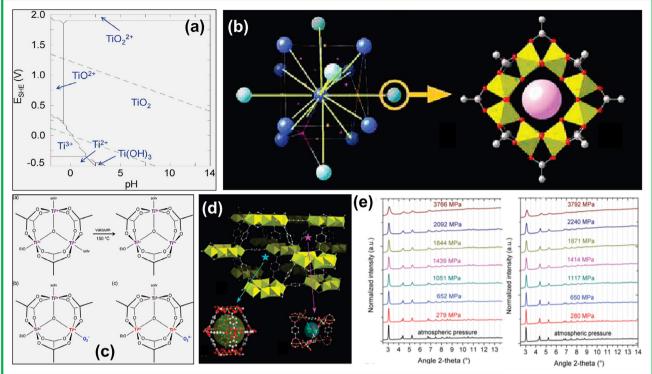


Fig. 1 (a) Pourbaix diagrams of titanium calculated for $[Ti^{4+}] = 10^{-3}$ mol L^{-1} at 25 °C using the Hydra and Medusa software. Green dashed lines: O_2/H_2O and H_2O/H_2 redox couples. Reprinted with permission from ref. 67. Copyright 2017 The Royal Society of Chemistry. (b) The perspective view of a centered cubic (cc) arrangement and the 12-fold coordination is shown by yellow lines. Purple and orange dots indicate the positions of the centers of the tetrahedral and octahedral vacancies and the view of the perforated cyclic octamer with edge- and corner-sharing Ti octahedra; it corresponds to the atom with an orange circle of the classical cc packing through the SBU augmentation. Reprinted with permission from ref. 47. Copyright 2009 American Chemical Society. (c) Illustration of the proposed composition of the Ti_3O clusters in 1 after heating at 150 °C under vacuum, reacting with O_2 to form titanium(IV) superoxide, and reacting with O_2 to form titanium(IV) peroxide. Reprinted with permission from ref. 69. Copyright 2015 American Chemical Society. (d) Perspective view of MIL-125 with the central octamer surrounded by 12 others; the pink and blue stars indicate the centers of the tetrahedral and octahedral vacancies in MIL-125, ball and stick representation of the octahedral vacancy, filled by water molecules (in green); the large yellow sphere represents the effective, accessible volume of the cage, and the tetrahedral vacancy. Reprinted with permission from ref. 47. Copyright 2009 American Chemical Society. (e) Powder X-ray diffraction patterns ($\lambda = 0.694120$ Å) were obtained for MIL-125(Ti) and NH₂-MIL-125(Ti) as a function of the applied pressure at room temperature. Reprinted with permission from ref. 70. Copyright 2015 The Royal Society of Chemistry.

form M–O–C clusters SBUs—is crucial for predicting topologies and structures. The coordination geometry can vary significantly with different reaction systems. Titanium oxo clusters with nuclearities ranging from 2 to 52 have been observed, reflecting the challenges in controlling Ti ion chemistry. Ti (IV) typically forms octahedrally coordinated TiO₆ units linked by μ_2 -O or μ_3 -O atoms, ensuing in variable Ti-oxo clusters. These clusters can vary even with a similar organic ligand under dissimilar synthetic environments. For instance, the Ti₈O₈(OH)₄ cluster in MIL-125 or NH₂-MIL-125 markedly differs from the Ti₃(μ_3 -O) cluster in its polymorph Ti-MIL-101. Such observations highlight the unpredictability and complexity associated with the working of Ti as a node in MOF structures.

One of the salient aspects of Ti-MOFs is their propensity to maintain very high symmetry, especially that of *in situ*-formed Ti-oxo clusters during solvothermal procedures. Most titanium cores, such as ${\rm Ti_8O_8(OH)_4}$ of MIL-125/NH₂-MIL-125 and ${\rm Ti_3(\mu_3-O)}$

of Ti-MIL-101/Ti-MIL-100, exhibit high symmetry, laying the foundation for extremely symmetrical MOFs with arranged pore structure or channels (Fig. 1b and c). ^{47,69} These symmetrical structures represent a unique advantage of Ti-MOFs, as they offer both stability and versatility that can be fine-tuned to specific applications.

4. Structure of NH₂-MIL-125

When Ti cations react with polydentate carboxylate acids, they form massive Ti-oxo-carboxylate clusters characterized by high coordination numbers and nuclearity, leading to the diverse arrangement of the resulting Ti-MOFs. These clusters form due to the incomplete hydrolysis and subsequent polycondensation of metal forerunners, with H₂O derived through esterification of the carboxylate entity (not a part of the reaction) with alcohols, that is established from reacted Ti-alkoxides (*in situ*)

or deliberately added to solvent ones. The resulting clusters are composed of TiO6 octahedra that are edge- and cornersharing, interconnected through carboxylate linkers. This mechanism suggests that cluster formation is key to understanding the unique structural properties and functions of titanium-based MOFs like NH2-MIL-125. The reactivity of titanium ions, their coordination preferences, and the behavior of precursor molecules all influence the resulting structure and material properties. This interplay defines the chemistry between titanium and organic ligands, determining the framework's strength and flexibility.

In 2009, Serre and colleagues reported the earliest porous crystalline carboxylate-based titanium-MOF, i.e., MIL-125, which was tuned via a quick solvothermal approach. 47 The MIL-125 structure comprised Ti₈O₈(OH)₄(COO)₁₂ octamers (cyclic), comprising edge- and corner-sharing TiO₅(OH) octahedra interconnected using ATA linkers (Fig. 1d). These SBUs were arranged in a centered cubic (cc) pattern, forming a quasi-cubic tetragonal lattice with an individual octamer having 12 adjacent SBUs. MIL-125(Ti) exhibited a rigid 3D interrelated porosity with two kinds of micropores, measuring 6 and 12.5 Å. Subsequently, the NH₂-functionalized correspondent NH₂-MIL-125 was efficaciously produced using the same linker, H₂ATA.⁷⁴ The inclusion of -NH₂ moieties in the organic ligands was unable to alter the MOF framework, and NH₂-MIL-125 was isostructural with the pristine MOF. The surface areas (BET & Langmuir) of NH2-MIL-125 were 1302 and 1719 m² g⁻¹, respectively, which were less than the theoretical value for the parent MIL-125(Ti) (approximately 2140 m² g⁻¹). This reduction was probably a result of partial pore obstruction by the free ligand. Introducing functional groups such as -NH₂ to the organic ligands improved the chemical stability of the material. NH₂-MIL-125 was more stable to humidity than MIL-125, owing to the intramolecular hydrogen bonds between the amino groups and carboxylate oxygen atoms. This added stability may extend its applications into environments with varying moisture and temperature.

Furthermore, Yot et al. explored the framework practices of MIL-125 and aminated-MIL-125 under external stress using high-pressure PXRD (3.5 GPa) (Fig. 1e).⁷⁰ The materials exhibited a gradual and reversible reduction in crystallinity when subjected to pressure. This type of MOF material can incorporate a high density of immobilized Ti sites within porous structures and form isostructural MOFs, whose photoresponsive efficiency can be readily adjusted by integrating BDC derivatives. The potential for NH2-MIL-125 to integrate a high density of immobilized Ti sites into its porous structure is significant for photocatalytic applications. The organic linkers can be adjusted to optimize the role played by titanium as an active site, thus allowing fine-tuning of Ti-MOFs for specific catalytic processes.

Synthesis procedures

The direct preparation of NH₂-MIL-125(Ti) is predominantly conducted through traditional hydrothermal or solvothermal

methods. This approach involves dissolving or dispersing organic ligands and Ti precursors in appropriate solvents, which are then placed in a PTFE-lined steel autoclave or a sealed Pyrex vessel and thermally treated under high pressure for a duration that can range from several hours to days.

Initially, hydrothermal techniques were used to develop a of titanophosphates^{75–77} and oxalatotitanates,⁷⁸ although most of these were nonporous materials. Among these, the first porous Ti-MOF, MIL-91,⁷⁷ was notable. Serre and his team successfully produced the renowned prototypical titanium-MOF, MIL-125, in 2009. This synthesis marked the creation of the first-ever titanium-carboxylate-based MOF with high crystallinity and porosity, accomplished through a solvothermal method. 47 In this synthesis, Ti(iOPr)4 reacted with terephthalic acid in a solvent blend of DMF and methanol at 423 K for 15 hours, resulting in a highly crystalline white powder identified as MIL-125(Ti). Perhaps the greatest drawback with MIL-125 is that it degrades fairly readily under extreme conditions (e.g., high moisture and temperature variation). There lies the potential for even newer, more stable constructions, just as easily accessible from the laboratory bench, with comparable or even superior durability. The NH2functionalization of the MIL-125(Ti), denoted as NH2-MIL-125 (Ti), is a good example of how minor modifications can improve its properties. It was synthesized by substituting terephthalic acid with 2-amino terephthalic acid in a comparable solvothermal procedure.74

The optimal selection of solvent is critical as it significantly impacts the crystallinity of the end products, often requiring a trial-and-error method to identify the optimal solvent combination. For example, although NH2-MIL-125 can be synthesized with DMF and methanol, the material's properties can vary dramatically depending on the precise solvent ratio and conditions applied. This brings out a point: although solvothermal synthesis is very effective for most purposes, it is less accurate and reproducible compared with newer, more controlled methods.

The reaction mixture's pH impacts the organic linker's deprotonation process and affects the metal ions' coordination geometry. The ideal pH range favors the development of wellorganized crystalline structures, while extreme pH values result in amorphous or less crystalline materials. Furthermore, the concentration of reactants is a critical factor in shaping the morphology of the resulting MOFs. For instance, regulating the concentration of the preliminary solution can change the morphological structure of NH2-MIL-125 microcrystals from spherical dishes to octahedra (Fig. 2a).79 The solvothermal reaction's length impacts the crystal development and the material's ultimate morphology. Prolonged reaction periods typically produce larger, well-structured crystals, while reduced time frames result in smaller, underdeveloped ones. This indicates that small changes in reactant concentrations and reaction conditions can significantly affect material properties, making careful optimization crucial.

While the solvothermal technique is the most predominant approach used for producing Ti-MOFs, optimizing conditions

(b)_{MIL-125} (c) g-1, STP) STP) NH₂-MIL-125(M) adsorbed (cm³ (E) 300 IVp/dDp 200 200 Pore diameter (Å) Pore diameter (Å 0.8 0.6 Relative pressure (P/P₀) Relative pressure (P/P₀) (e) ₁₂₀ (f) 80 $V_{ads} / cm^3 g^{-1}$ **UV** irradiation (a) 0.6 0.8 1.0 500 600 700 800 200 300 400 HCOO (g) 8 Light HCOO_] / hmol visible light TEOA NH₂-MIL-125(Ti) NH₂-MIL-125(Ti)

Fig. 2 (a) SEM images of NH₂-MIL-125(Ti) crystals synthesized with different total solvent volumes: 40 mL, 30 mL, 20 mL, and 15 mL. Reprinted with permission from ref. 79. Copyright 2014 The Royal Society of Chemistry. (b) N₂ adsorption—desorption isotherms at 77 K and (c) HK pore size distribution curves of MIL-125, NH₂-MIL-125, and NH₂-MIL-125(M) samples. Reprinted with permission from ref. 80. Copyright 2012 Elsevier B.V. (d) View of MIL-125 in a capillary filled with benzyl alcohol before and after UV irradiation. Reprinted with permission from ref. 47. Copyright 2009 American Chemical Society. (e) UV/Vis spectra of MIL-125(Ti) and NH₂-MIL-125(Ti). The inset shows the samples. (f) CO₂ adsorption isotherms (1 atm, 273 K) of MIL-125(Ti) and NH₂-MIL-125(Ti). (g) The amount of HCOO⁻ produced as a function of the time of irradiation over NH₂-MIL-125(Ti) (\square), a mixture of TiO₂ and H₂ATA (O), and visible light irradiation without a sample (\triangle). (h) Proposed mechanism for the photocatalytic CO₂ reduction over NH₂-MIL-125(Ti) under visible light irradiation. Reprinted with permission from ref. 50. Copyright 2012 Wiley-VCH Verlag GmbH & Co.

for large-scale production remains a challenge. Hence, in addition to the conventional solvothermal approach, alternative technologies such as microwave^{80,81} and ultrasonic⁸² procedures have been employed for the facile fabrication of Ti-MOFs with high porosity. These techniques accelerate the process of crystallization, leading to reduced reaction times and temperatures. Ti-MOFs synthesized using these alternative approaches often exhibit varied particle shapes and sizes, which can significantly impact their characteristics. For instance, the rapid heating capability of microwave procedures was first used to formulate NH₂-MIL-125 through a solvothermal scheme.⁸⁰ This approach decreased the reaction time from 16 hours, typical in traditional solvothermal methods, to just 1 hour using microwave heating. PXRD confirmed that the

resulting powders were phase-pure, with a typical particle size of 0.7 mm, which was less than that obtained via the solvothermal method. The lowering in particle size was attributed to the shortened reaction duration limiting particle growth and the speedy nucleation powered by the thermal process. The BET surface area of NH₂-MIL-125(M) was determined to be 1492 m² g⁻¹, with a pore volume of 0.66 cm³ g⁻¹, compared with 1469 m² g⁻¹ and 0.65 cm³ g⁻¹, respectively, for NH₂-MIL-125 (Fig. 2b and c). The smaller particle sizes indicated for these samples could mean that microwave synthesis might provide more control of particle size, but increased defect density or less structural integrity may also come with some downsides. In this regard, there is often a balance to strike between speed and material quality.

Ultrasonication generates and disintegrates bubbles within the solvent, creating localized high temperatures and pressures that lead to rapid heating and cooling. This environment promotes fine crystallization while intense shear forces excite molecules, accelerating precursor dispersion, bond formation, and breakage. Han et al. found that the Ti-based NH2-MIL-125 prepared using an ultrasonication-assisted solvothermal approach displayed an even distribution in sizes of the nanoparticle, approximately 300 nm.82 Additionally, a regular and defect-free MOF membrane was successfully created on a porous TiO2 disk using a combined microwave/ultrasonic/ solvothermal process.⁸³ Attempts to synthesize NH₂-MIL-125 (Ti) crystals on the support using microwave (MW) or ultrasonication (US) alone failed, highlighting the need for a seeding step for heterogeneous nucleation. MW irradiation was more effective in initiating nucleation than the US, and subsequent US treatment densified the crystal population using pre-formed crystals as secondary nucleation sites. The resulting NH₂-MIL125(Ti) membrane showed densely packed, uniform crystals without impurities, confirmed by SEM, XRD, N₂ adsorption-desorption isotherms, and EDS analysis. The membrane exhibited high surface area, pore volume, and strong adhesion to the TiO2 support. Its heterogeneous catalytic activity was evaluated for the Knoevenagel condensation reaction, suggesting its potential as a recyclable catalyst. The combined MW and US approach effectively produced a highquality NH2-MIL-125(Ti) membrane with enhanced catalytic performance. The success of hybrid approaches thus raises a key question: synthesis complexity vs. quality of materials. Combining multiple techniques might lead to some improved properties; however, this would also increase the complexity of synthesis.

Our investigation revealed that tetra butyl titanate and titanium isopropoxide are the primary titanium sources for synthesizing NH₂-MIL-125(Ti), paired with 2-amino terephthalic acid as a consistent raw material. Adjusting the proportions of these sources yields varied morphologies without a discernible pattern. Solvents like DMF and CH3OH are commonly used, typically in a 9:1 volume ratio, which often produces tetragonal nanosheets, nano disc cakes, and decahedral prisms. Reactions are generally conducted at 150 °C for 16-72 hours, with 24 hours being the most common. Morphologies-from disc cakes to irregular nanosheets and octahedrons-are influenced by reaction time, pH changes, and the titanium source ratio to 2-amino terephthalic acid, while reactor size has minimal impact. Surfactants and organic reagents, such as CTAB and ascorbic acid, are often added to achieve precise morphology control. Direct synthesis of NH₂-MIL-125(Ti) through hydrothermal or solvothermal methods suffers from a lack of optimization of solvents and reactant concentrations for crystallinity and morphology. Other methods, such as microwave and ultrasonic synthesis, improve crystallization, shorten reaction times, and control particle sizes. Future efforts should target consistency and scalability. While microwave and ultrasound methods press forward, however, issues around reproducibility and largescale implementation remain. Most work has focused on the crystallinity and surface area of NH2-MIL-125(Ti), whereas less attention has been paid to stability under real-world conditions. Long-term stability, especially in catalytic and industrial applications, needs to be the focus of attention. Besides, it is of utmost importance to examine green solvents for the synthesis of NH2-MIL125(Ti) since the current solvothermal routes mostly work with toxic solvents, and new sustainable alternatives should be investigated to see their impact on both synthesis efficiency and material properties.

Significance of -NH₂ functionalization

The incorporation of amine groups (NH₂) into MOFs has come forth as an effective strategy to improve both their stability and their response to visible light. The existence of amine groups alters the electronic structure by generating additional energy levels within the framework, leading to a decrease in the overall band gap of the material. These changes facilitate charge transfer through alternative pathways and broaden the MOF's absorption spectrum into the visible and near-infrared regions, thereby enhancing its photophysical properties.84 The amine group, rich in electrons due to the lone pair on the nitrogen atom, can elevate the Fermi level of material, shifting the CB position and thereby lowering the energy barrier for charge transfer. This effect enhances the photogenerated charge carrier separation. Additionally, the -NH2 group facilitates charge transference from the organic linker to metal-oxo clusters, promoting their excitation.85 Due to the Lewis acidity of amines, they also contribute to an increased generation of free electrons by effectively injecting them into the conduction band of semiconductors. The lone pair of the amine group takes part in extended conjugation within the MOF assembly, which lowers the rate of recombination of charge carriers.86 Amine groups can also be charge carriers as they trap photogenerated electrons, which can either delay recombination or act as active sites for reactions with adsorbed species. In addition, amine groups modify the electronic properties of the ligand, enhancing its ability to interact with the metal center. This affects the electronic transitions, with optical absorption being transitioned from the UV region to the visible region. The lowering of the excitation energy arises from improved π -conjugation as well as from the electron-donating effect by the amine groups.⁸⁷ Amine-functionalization could also induce the formation of localized electronic states responsible for visible-light absorption. These states can modify the material's optical properties, thereby facilitating visible-light-driven photocatalysis by enabling the absorption of lower-energy photons.

To enhance the MOF stability and reliability, the metal center is coordinated with the -NH2 group of organic ligands, forming robust metal-NH2 bonds. These bonds provide additional structural stability, protecting the MOF from changes during photocatalytic reactions. The amine groups also act as a protective measure against pH fluctuations Review Nanoscale

induced by reactive species generated during photocatalysis. Furthermore, amine groups help prevent oxidative degradation of the MOF by scavenging reactive oxygen species (ROS), such as hydroxyl radicals ('OH) and singlet oxygen (1O2), thereby diminishing their concentration. When ROS are present, the -NH₂ groups experience oxidation, donating electrons to protect the MOF's ligands from oxidation and preserving the material's architectural integrity.⁸⁸ Water is frequently taken as a reactant in photocatalysis, although some MOFs may exhibit sensitivity to moisture. Under photocatalytic circumstances, -NH₂ functionalization can enhance the aqueous stability of MOFs by incorporating either hydrophobic or hydrophilic properties based on the specific type of -NH2 groups incorporated. 89-91 The upgraded photocatalytic performance is a result of several factors, including enhanced visible light uptake, efficient excitons' separation and transfer, selective interactions with reactant species, the formation of reactive sites, and increased chemical stability. Collectively, these factors contribute to the significant and beneficial impact of MOFs with NH₂-functionalization on photocatalysis.

7. Advantages of NH₂-MIL-125 as photocatalyst

NH₂-MIL-125(Ti) is a titanium-based metal-organic framework (MOF) featuring amine-functionalized organic linkers. This

structure offers significant potential in photocatalysis, attributed to its multiple key advantages, as displayed in Scheme 2.

7.1. Improved visible light uptake

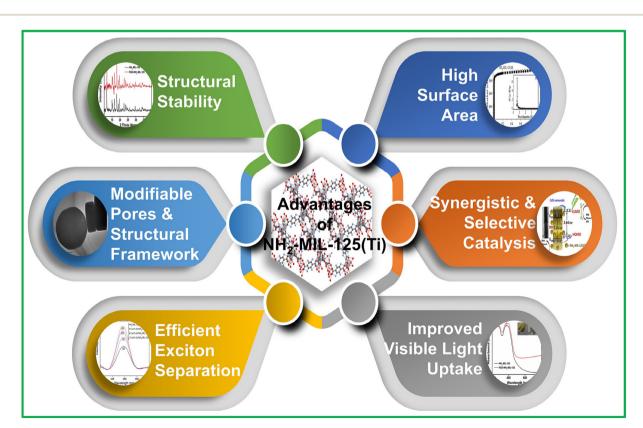
The $-\mathrm{NH}_2$ functionalization in NH_2 -MIL-125(Ti) extends light absorption into the visible range, enabling efficient photocatalysis under natural sunlight. This material displays heightened molar absorptivity owing to the $-\mathrm{NH}_2$ group and the organic linker, effectually exploiting the incident light energy. The synergistic interaction between the Ti center and the NH_2 -functionalized ligand further enhances light absorption. 92

7.2. Efficient exciton separation

In NH_2 -MIL-125(Ti), the incorporation of $-NH_2$ groups into the organic linkers enhances photocatalytic efficiency by promoting charge separation and suppressing e^--h^+ recombination. Acting as electron donors, $-NH_2$ groups create localized electronic states, adjust the band structure to reduce the band gap, and enable greater visible light uptake. The well-ordered crystal structure, with robust association of Ti nodes and NH_2 -functionalized ligands, ensures efficient electron transport and effective charge transfer to catalytic sites. 93

7.3. Modifiable pores and structural framework

The tunable pore structure and structural framework of NH₂-MIL-125(Ti) offer significant advantages for photocatalysis, including improved reactant diffusion, faster reaction kinetics,



Scheme 2 Advantages of NH₂-MIL-125(Ti)-based photocatalysts.

and enhanced efficiency. The optimized surface area supports selective molecular adsorption and increases active site availability, promoting better light absorption. The tailored pores enable spatial separation of photoinduced e⁻ and h⁺, boosting charge carrier efficacy. Additionally, the controlled pore architecture manages reaction intermediates effectively, reducing side reactions and improving selectivity. These features make NH₂-MIL-125(Ti) adaptable for numerous photocatalytic applications, ranging from environmental remediation to energy harvesting and chemical synthesis, delivering high efficiency and versatility.94

7.4. Structural stability

NH₂-MIL-125(Ti) exhibits remarkable thermal and chemical robustness, confirming its structural integrity and effectiveness across a range of reaction conditions. This robustness supports long-term photocatalytic activity without significant degradation, making it highly reliable for prolonged use. Its durability in harsh chemical conditions, encompassing both acidic and alkaline environments, broadens its applicability and ensures sustained functionality in demanding settings. The stable structure protects active catalytic sites, reducing the likelihood of metal ion or linker leaching, thereby preserving the reaction medium's purity and maintaining high photocatalytic efficiency. These characteristics enhance its durability, reliability, and cost-effectiveness, making it a viable option for diverse environmental and industrial processes. 95

7.5. High surface area

NH2-MIL-125(Ti) has a significantly large surface area, enhancing photocatalytic efficiency by facilitating the formation of active sites, improving reactant adsorption, optimizing light harvesting, and facilitating exciton pair employment. Its porous structure supports efficient reactant diffusion and adsorption, while surface functionalization further expands its scalability for practical uses. These features synergistically boost reaction rates and overall performance, positioning it as a potential material for addressing energy and environmental challenges. 96,97

7.6. Selective catalysis

NH₂-MIL-125(Ti) stands out as a highly adaptable photocatalyst, leveraging its amine-functionalized groups and tunable pore structure for enhanced performance. The -NH2 groups and tailored pores enable selective interactions with reactant molecules, allowing adsorption based on size, shape, and functional groups. This selectivity increases the localized abundance of reactants at active sites, accelerating reaction rates and boosting catalytic efficiency. The amine groups also create specific reaction sites that promote precise engagements with target molecules, improving the output and selectivity of target products. The modifiable pore configuration provides a controlled microenvironment, allowing fine-tuning of reaction conditions for optimized performance. This structure supports the selective reduction or oxidation of compounds based on reaction conditions and target molecules, enhancing its versatility. Applications include CO2 conversion into value-added chemicals and the selective oxidation of organic contaminants, demonstrating the material's potential for energy harvesting and environmental mitigation. 98,99

7.7. Synergistic effects

The interaction between the metal node and ligands in NH₂-MIL-125(Ti) creates a synergistic impact that boosts its photocatalytic performance. This synergy broadens the light absorption into the visible range, reduces charge recombination, and accelerates reaction kinetics by stabilizing intermediates and lowering activation energy. It also promotes the generation of reactive oxygen species, aiding in the efficient decomposition of organic pollutants. The durable metal-ligand framework ensures stability and long-term usability, making it ideal for industrial and environmental applications. Moreover, the material's design can be customized by modifying the metal center or ligands to suit various needs, including environmental cleanup, renewable energy generation, and chemical production. 100,101

NH₂-MIL-125(Ti) offers notable benefits in photocatalysis, including enhanced visible light uptake, efficient exciton separation, extensive surface area, selective adsorption, and robust framework stability. These attributes make it suitable for diverse applications, for instance, water splitting, pollutant degradation, CO2 reduction, and organic transformations. However, challenges like complex synthesis, potential degradation, restricted stability in aqueous conditions, resource demands, by-product management, and scalability must be taken into account to unlock its full strength in practical uses.

Photocatalysis in NH₂-MIL-125 8.

Ligand-to-metal charge transfer (LMCT) is the most general method for achieving exciton pair separation in MOFs. The LMCT pathway typically has a low rate of e transfer and a high tendency of recombination of charge carriers, which can be upgraded by fine-tuning linker-metal cluster interaction or altering the charge transportation way to enhance photocatalytic performance. The first ever Ti-MOF to exhibit substantial photoresponsiveness was the classic MIL-125, along with its NH₂-functionalized analog, NH₂-MIL-125. Being the first Ti-carboxylate-based MOF, the finding of MIL-125 undoubtedly triggered the current surge in Ti-MOF research, spurring the advancement of newfangled Ti-MOF structures and applications across numerous fields. Notably, MIL-125(Ti) exhibits a fascinating color transformation from white to purple-gray-blue under UV radiation in the presence of alcohol (Fig. 2d).47 This photochromic behavior indicates photocatalytic activity; however, MIL-125(Ti) can undergo excitation only by UV radiation. In order to expand the light absorption to the visible range, Li et al. utilized NH2-BDC in place of H₂BDC in MIL-125(Ti), resulting in the NH₂-functionalized NH₂-MIL-125(Ti).⁵⁰ The thoughtful introduction of amine functionalization in NH2-MIL-125(Ti) introduced extended

light harvesting in the visible range along with enhancement in CO_2 adsorption ability (Fig. 2e and f). With triethanolamine as a sacrificial agent, the MOF CO_2 can be reduced to formate in visible light (Fig. 2g). ESR study showed that the ATA ligand performed as a light probe, harvesting light and becoming excited, following the e-transportation to the titanium-oxo clusters to generate Ti^{3+} for CO_2 reduction (Fig. 2h). This effective display of photocatalytic CO_2 reduction using NH₂-MIL-125(Ti) marked the very first instance of a visible-light-induced CO_2 reduction reaction utilizing MOFs, opening new perceptions for developing photocatalysts valuable for CO_2 reduction.

The LMCT mechanism in NH₂-MIL-125(Ti) was further confirmed by García et al. through their study of the transient states generated upon photoexcitation. 102 The article provides strong evidence for photoinduced charge separation in NH2-MIL-125. Laser flash photolysis of dry powder and aqueous suspensions showed a transient spectrum with broad absorption (380-800 nm), indicating charge separation upon excitation. The transient signal, with long lifetimes, decreased in the presence of oxygen, suggesting quenching of photogenerated species. Blue-colored radical ions of the electron acceptor viologen) and donor (N,N,N',N'-tetramethyl-p-(methyl phenylenediamine) confirmed NH2-MIL-125's ability to drive reduction and oxidation reactions. Comparison with TiO2 (P25) showed similar photochemical behavior, reinforcing NH2-MIL-125's semiconductor nature. Despite challenges such as slow electron transfer and high exciton recombination via the LMCT pathway, incorporating functional groups like amines broadens the light absorption spectrum into the visible range and enhances charge separation. NH2-MIL-125 (Ti) is thus appropriate for photocatalytic implementations in energy generation, contaminant degradation, and organic synthesis. Even though NH₂-MIL-125(Ti) is quite promising, fundamental issues, such as charge transport and recombination, hamper its photocatalytic efficiency. These cannot be solely overcome by functional group modification, but rather it requires further mechanistic insights combined with the development of novel strategies that could improve electron transport as well as integration of MOFs with other materials to improve catalytic properties.

9. Improvement strategies for NH₂-MIL-125-based materials

Photocatalysis typically includes three primary phases: (i) light absorption to produce a charge-separated state, (ii) separation and relocation of these exciton pairs, and (iii) origination of redox reactions via photoinduced excitons. ¹⁰³ Effective photocatalysis requires a photocatalyst with strong light absorption and rapid exciton pair separation and migration. Flexible frameworks of MOFs permit modifications in structures that compel the addition of titanium ions for the creation of photoactive sites, followed by elongation of π -conjugated linkers to improve charge separation and transport. In titanium-based

MOFs, enhancing photocatalytic performance involves crystal engineering of metal nodes, ligands, and voids. Strategies like post-synthetic modification, linker functionalization, and metal node doping are used to improve optoelectrical properties and exciton pair separation (Fig. 3a). ¹⁰⁴

9.1. Linker modification

Several reports have shown that substituents such as -NH₂, -NO2, -OH, and -SH on MOF ligands significantly influence the $E_{\rm s}$ of the MOFs. ^{108–110} Specifically, –NH₂ functionalization is an efficient technique to shift the optical absorption capacity of MOFs toward the visible-light region. 47,111 Given that most of the sunlight energy is in the Vis-NIR range, piloting the photochemical activity towards the visible range is essential, which calls for E_g alteration of the concerned photocatalysts. MIL-125(Ti) has an optical E_g value of approximately 3.6 eV with an absorption maximum of 345 nm, suggesting its photoactivity under UV irradiation. For Ti-based MOF conamino-modified, photosensitizer-functionalized ligands, electron-rich-conjugated, and with extended light-harvesting range are time and again chosen. 47 NH2-MIL-125 was the first example, where introducing an NH2-group into the ATA linker matrix produced an additional absorption maximum at 550 nm (visible light region), decreasing the $E_{\rm g}$ value to 2.6 eV, giving the MOFs noteworthy photocatalytic activity for CO2 reduction. 50 This functionalization is practically appealing because it provides the photocatalysts with extended functional wavelengths—giving photocatalysts access to a broader range of sunlight, which is important in largescale application scenarios.

To elucidate the exact influence of -NH2 on the optical characteristics of MOFs, Walsh et al. conducted a joint computational and experimental analysis (Fig. 3b). 105 They studied MIL-125(Ti) analogs with varied amounts of -NH2 sites per unit cell (approximately 1, 6, and 12 BDC-NH2 linkers) and found that the samples exhibited the same E_g value of nearly 2.6 eV, with the molar extinction coefficient rising proportionately with the ATA content, indicating that a solo -NH₂ group is accountable for the decreased $E_{\rm g}$ in the mono-aminated MIL-125(Ti). Electronic environment experiments indicated that adding -NH₂ to MIL-125(Ti) elevates the VB by 1.2 eV, leaving the CB unaffected. Adding another -NH2 group to the ligand further marked a decreased $E_{\rm g}$ value. Assessment of the substituents with varied e donating abilities, such as -CH₃, -OH, and -Cl, revealed the fact that strong e donating substituents notably reduced the E_g , which justifies the most significant redshift exhibited by di-aminated BDC-(NH₂)₂ linker. This study depicts the function of -NH₂ in the optical attributes of MOFs and guides the design of MOFs with excellent light-harvesting properties. Van Der Veen and co-workers' recent studies demonstrated that besides enhanced light absorption, -NH₂ in NH₂-MIL-125(Ti) also performs as a hole stabilizer, prolonging the photoexcited state's life period and allowing effective utilization of photogenerated charges for photocatalytic reactions. 112 Hendon and his team used experimental and computational methods to explore the influence of the

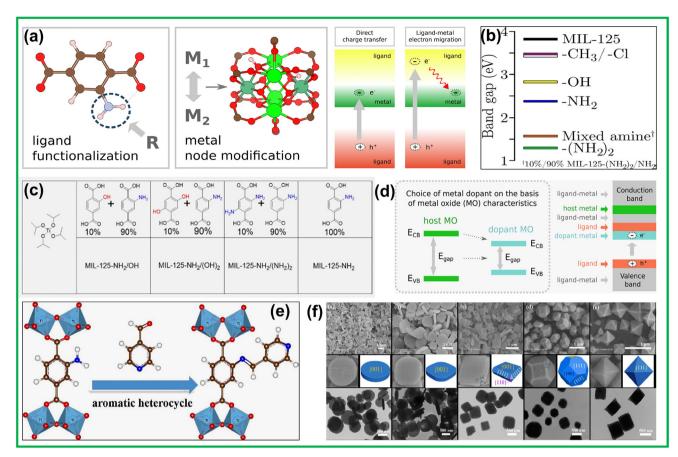


Fig. 3 (a) Strategies for compositional tuning of MOF photoactivity by ligand functionalization or metal node modification. Excited state pathways to achieve a metal node reduction in a MOF. Reprinted with permission from ref. 104. Copyright 2019 American Chemical Society. (b) HSE06-predicted band gaps of MIL-125 (black) and its analogs containing functionalized BDC linkers. Reprinted with permission from ref. 105. Copyright 2013 American Chemical Society. (c) Tabular representation of NH2-MIL-125/OH, NH2-MIL-125/(OH)2, NH2-MIL-125/(NH2)2, and NH2-MIL-125, made of titanium isopropoxide and either a 1:9 ligand ratio or only a NH2-BDC ligand. Reprinted with permission from ref. 106. Copyright 2021 American Chemical Society. (d) Modification of band gap, CB alignment, and creation of localized electron traps using metal doping of a MOF. Reprinted with permission from ref. 104. Copyright 2019 American Chemical Society. (e) Post-synthetic strategy of grafting NH2-MIL-125(Ti) with aromatic heterocycles. Reprinted with permission from ref. 51. Copyright 2017 Elsevier B.V. (f) SEM, enlarged SEM, TEM images, and the corresponding 3D geometry models of as-synthesized NH2-MIL-125(Ti) obtained under different concentrations of CTAB. Reprinted with permission from ref. 107. Copyright 2019 The Royal Society of Chemistry.

functional modification of ligands on the E_g in depth. They found that bringing in a di-aminated ligand further lowered the E_g compared with NH₂-MIL-125. The synthesis of mixedamine MIL-125, which had 10% and 90% BDC-NH2 ATA ligands, achieved an experimental $E_{\rm g}$ value of 1.3 eV that showcased a significant shift towards the red/IR region. The -NH2 group does enhance electron density and light absorption, but it also makes the material more susceptible to thermal and photoinduced instability owing to the oxidatively unstable nature of amines, especially under prolonged exposure to UV and visible light. These issues are not fully treated within the current literature. In particular, the practically important aspects of real-world applications, such as long-range stability and scalability, need to be explored in greater depth, especially concerning amine-functionalized MOFs.

He et al. observed that introducing defects into the ligand structure via a one-step heating-stirring method reduced the E_g

from 2.70 to 2.60 eV and created both microporous and mesoporous structures, which exposed more metal active sites. The inclusion of defects in the original MOF matrix expanded visible light harvesting and promoted the full utilization of solar energy.¹¹³ Defects may often create localized electronic states that, if not carefully controlled, can become trap states, increasing charge carrier recombination or giving rise to unwanted side reactions. This is an ever-recurring problem with many photocatalytic materials, putting into question the controllability of defect density. Overengineering defects may enhance charge separation in photocatalysis but may instead lower structural integrity or photochemical stability at the same time.

Similarly, Mohammadnezhad et al. explored the optical characteristics of NH₂-MIL-125 MOF by incorporating three variants of functionalized BDC ligands - (NH₂)₂-BDC, (OH)₂-BDC, and OH-BDC - into the framework at relatively low con-

centrations (Fig. 3c). 106 Incorporating multiple ligands with varying functionalization into the MOF matrix is an alluring tactic to enhance the efficiency of the photocatalyst. 47 The method yields a potential photocatalyst owing to the π-conjugation formed by individual chromophore-functionalized units coupled loosely. 105,114 With chromophoric -NH2 and/or -OH groups as the visible light-harvesting sites, it promotes an e⁻ from the VB maxima to the CB minima. This transition in NH₂-MIL-125 takes place from the π -orbital of the ligands to the oxygen's 2p orbital along with the titanium's split 3d orbital. Amid the process, exciton pairs are created, with h⁺ confined to -NH₂ and e⁻ at relative liberty to transit from the linker to the metal cluster, stabilizing the photoexcited charge carriers that allow e migration to the MOF's exterior. The -OH and -NH2 functional groups, chosen for their e-donating capability, were located on the ATA linker because of the presence of free electron pairs that do not take part in coordination, which enhances LMCT increasing both the overall reaction rate as well as the e- density on the ligand. 115 Increasing electron density enhances electron delocalization upon radiation, boosting photocatalytic efficiency. The chromophoric -NH2 and -OH groups make the catalyst active in the visible spectrum, giving the material orange or yellow hues and improving light absorption. NH₂-MIL-125/ (OH)₂ shows a redshift with stronger absorption, partly due to $p-\pi$ conjugation between the -OH groups and the ATA linker's benzene ring. While NH₂-MIL-125 and NH₂-MIL-125/OH have comparable Eg values of approximately 2.6 eV, NH2-MIL-125/ $(OH)_2$ has a lower E_g of 2.2 eV, enhancing visible light uptake. Although NH_2 -MIL-125/ $(NH_2)_2$ exhibits the most significant

Logan et al. created a series of NH2-MIL-125 analogs by switching the NH₂-groups with alkyl groups of varying chain lengths and connectivity. 116 The incorporation of these functional groups led to a reduction in the $E_{\rm g}$ value, decreasing from 2.46 eV to 2.29 eV for MIL-125-NHMe and for MIL-125-NHCyp, respectively. MIL-125-NHCyp, featuring a small E_g , an extended excited state lifetime of 68.8 ns, and a high AQY of 1.80%, exhibited significantly enhanced photoactivity compared with the original MOF. These findings indicate that the optoelectronic properties of NH2-MIL-125 MOF-based materials can be effectively modulated through the strategic functionalization of the linking units. In conclusion, while the strategies listed above may show great promise, we believe there is a need for this field to broaden its scope from the mere demonstration of enhanced photocatalytic activity in a controlled laboratory to tackle practical dilemmas. Future works must also address multi-strategy optimization that adopts functions like functionalization, defect engineering, alkyl substitution, etc., giving rise to more holistic approaches that assume real-world conditions.

redshift (E_g = 1.6 eV), its absorption intensity is weaker than

NH₂-MIL-125/(OH)₂. PL spectra at 420 nm show similar emis-

sion peaks near 450 nm for all samples. However, NH_2 -MIL-125/(OH)₂ exhibits a redshift to ~510 nm, highlighting

the energy gap shift toward the visible spectrum with (OH)2-

9.2. Metal node modification

The modification of metal nodes in MOFs is one of the most powerful strategies for enhancing photocatalytic efficiency. The inherent advantage of MOFs, including NH₂-MIL-125, lies in their ability to host coordinatively unsaturated metal sites (CUSs). MOFs often feature metal nodes initially coordinated with solvent molecules such as DMF or H₂O. By applying heat or using vacuum activation, these solvent molecules can be eliminated, leading to the formation of CUSs. These sites enable MOFs to engage in Lewis acid catalytic reactions.¹¹⁷ These CUSs, however, are masked or deactivated by organic ligands coordinating with metal centers. This, in turn, limits the chances of successful charge transfer and catalytic performance.¹¹⁸ Enhancing the photocatalytic efficacy of MILs necessitates modifying the metal nodes to yield more efficient catalytic sites.

Zhang et al. illustrated that partly substituting Ti with Ru in NH₂-MIL-125 leads to increased exposure of titanium-oxo clusters, thereby improving its catalytic properties. 119 Replacing metal nodes partially, or by metal doping, within the SBUs represents a potential but comparatively underexplored approach to tailoring MOF characteristics. This method enables the adjustment of both the relative and absolute energy positions of metal nodes within a MOF's band structure. When choosing a metal for doping, it's essential to consider its electron affinity, which acts as a critical condition in defining the position of its states. Since metal nodes within MOFs can be conceptualized as fixed metal-oxide clusters, the reducibility measures of metal oxides can guide the selection of suitable dopants. Intended for MOFs, criteria such as a suitably small $E_{\rm g}$ value and superior e⁻ affinity, determined via the absolute position of the CB edge, are pertinent (Fig. 3d). 104,120,121 These metal nodes in MOFs, which bond with oxygen atoms from the polycarboxylate-ligands to create metal-oxo clusters, frequently act as catalytic active sites in photocatalytic processes. Titanium, having multi-valence states (Ti⁴⁺/Ti³⁺), emerges as a favorable choice for node incorporation in MOF photocatalysts due to its abundance in Earth's crust and non-toxicity. To facilitate efficient charge transfer, a MOF ought to have an optimum E_g value, with energetically favorable positioning of the CB and VB edges for photocatalytic reactions. A feasible approach for achieving this ideal E_g value involves the partial substitution of the metal nodes, producing mixed-metal center MOFs. This synthetic approach allows for fine-tuning the energy positions within the MOF band structure.

Given NH₂-MIL-125, metal doping aims to improve the localization of metal nodes while diminishing the input from ligands. For instance, Syzgantseva *et al.*, through a combination of computational simulations and experimental studies, demonstrated that incorporating V into NH₂-MIL-125 induces a notable redshift in the UV-Vis spectrum, changing its color from yellow to brown. This observation was supported by theoretical computations showing that V incorporation decreases the position of the CB edge. ¹⁰⁴ Doping NH₂-MIL-125 with W and Nb lowers its CB edge. Nb enhances metal node

BDC inclusion.

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localization, though the DOS still shows contributions from C atoms and Ti to the LUCO state. Nb_2O_5 has a 0.2 eV higher E_g than TiO_2 , with modest localization. WO_3 , despite its higher electron affinity, shows contributions from both Ti and W at the CB bottom, suggesting electron affinity alone doesn't guarantee localization. W-doping modestly reduces the CB position, aligning with optical E_g reductions in W-doped TiO_2 . For effective exciton separation, the dopant must create a more reducible metal oxide, lowering both E_g and the CB edge position. While these modifications enhance light absorption and charge transport, metal doping can occasionally result in unintended effects, like the creation of defective sites or unfavor-

able bonding configurations. The reducibility of the doped

metals, along with the risk of leaching or deactivation over

time, may undermine the overall catalytic performance. Li et al. produced a bifunctional photocatalyst, Yb-NH₂-MIL-125(Ti), coordinated with Yb, which plays a pivotal role in photocatalysis. During the reaction, charge carriers generated by light transfer from the organic ligand to the metal center instead of the Ti-O cluster, potentially through a transient Yb (II/III) center. This mechanism enhanced the rate of e-transportation and reduced the recombination of exciton pairs. 122 Compared with the neat material, the light absorption efficiency and absorption region of Yb-NM-X were improved, with a decrease in E_{σ} from 2.51 to 2.37 eV, indicating enhanced light energy utilization with the introduction of Yb species. In their research, Chen et al. opted for Ni, recognized for its superior affinity for e-s, as a dopant to inspect its impact on e-allocation within NH2-MIL-125. Their findings demonstrated that the inclusion of Ni²⁺ ions enhanced e⁻ transference and promoted carrier separation by modifying the electronic band structure of the Ti-O clusters in NH2-MIL-125. Due to Ni's greater electronegativity relative to Ti, its integration around the oxygen atoms of these Ti-O clusters resulted in a significant alteration of their electronic configuration. Ni doping favors the excitons' separation; on the contrary, W and Nb doping do not necessarily favor charge localization, which proves that electron affinity alone cannot render the process of doping successful. The energy alignment between the dopant and host material becomes a significant factor. It should not be only the case of high electron-affinity metals; doping strategies must also tune carefully the electronic structure. Studies show that there were only slight changes attributed to Nb and W doping regarding band gap; hence band gap narrowing does not always improve photocatalytic activity. The relationship between dopant concentration, localization of metal sites, and electronic structure is complex. Future work should be directed toward multi-doped materials, which synergistically integrate dopants but remain stable.

9.3. Encapsulation of active species

The combination of metal nodes and functional ligands in MOFs creates open frameworks with large voids and extensive internal specific surface areas. This structure allows MOFs to house active molecular catalysts individually, preventing agglomeration that could diminish catalytic performance.

Gascon et al. employed a "ship-in-a-bottle" synthetic approach to encapsulate a Co-dioxime-diimine catalyst inside the void of the photoactive NH₂-MIL-125. This resulted in a sturdy, multioperational Co@MOF photocatalyst, which was highly effective for photocatalytic hydrogen production in visible regions. 123 ESR studies demonstrated that the incorporation of the Co complex within the NH2-MIL-125(Ti) cavity facilitated effective charge transference from the excited MOF to the active Co(II) sites, greatly enhancing photocatalytic activity. As a result, a 20-fold increase in the hydrogen production of Co@MOF was observed compared with that of the parent NH2-MIL-125. Building on this strategy, Jiang et al. incorporated a Co(II) molecular catalyst, [CoII(TPA)Cl][Cl], into NH₂-MIL-125. 124 This composite enhanced photocatalytic hydrogen production efficacy by facilitating photogenerated charge transference from the MOF to the Co(II) complex. The incorporation of photon absorption, charge dissociation, electron movement, and catalytic activity within the guest@MOF multifunctional framework highlights the significant potential of MOFs as podiums for crafting versatile photocatalysts. This process has shown significant improvements in charge transfer and prevents catalyst leakage, a common occurrence in conventional methods. The drawback lies especially in the unexplored complexity of charge transfer mechanisms that exist among all such systems. Although the studies mentioned previously show the possible transfer of charges from the MOF to the enclosed catalyst, there would still exist some bottlenecks which could primarily include kinetic limitations and electron transport between the MOF matrix and the active species. It could, therefore, add more understanding to studies on charge dynamics within encapsulated systems.

Enhancing the reactivity of catalytic active centers is an additional approach for boosting photocatalytic activity, aiming to increase overall photochemical productivity. Noblemetal NPs such as Pt or Ru are highly efficient multifunctional catalysts widely recognized for their effectiveness in various critical reactions. Encapsulating nanoparticles like Pt, Pd, Au, or Ru inside the voids of NH2-MIL-125 has been documented and utilized to carry out photocatalytic processes. 56,125 In these structures, the separation of photoinduced exciton pairs is efficiently achieved, with e's migrating to the surfaces of the noble metal NPs rather than to the Ti-containing secondary building units. This electron transfer facilitates reduction reactions. The photocatalytic efficiency of coordinatively unsaturated Pt or Ru atoms far surpasses that of intrinsic Ti³⁺ sites in titanium-MOFs, resulting in enhanced overall photocatalytic performance. Additionally, incorporating other functional materials such as MoS2, hetero-poly acids, QDs, and carbon nano species with Ti-MOFs presents promising opportunities for further investigation. Encapsulating active species in MOFs is challenging due to issues like catalyst aggregation and uneven distribution. The "ship-in-a-bottle" approach improves photocatalytic activity but struggles with homogeneity and efficient charge transfer. Incorporating noble metal nanoparticles enhances activity through better charge separation and electron transport, but optimizing their uniform distri-

bution and interaction with the MOF matrix is essential. Exploring the loading of other functional materials and studying charge transfer mechanisms will be key to developing more effective and stable photocatalysts.

9.4. Post-synthetic modification

Post-synthetic modification (PSM) is a crucial technique for the functionalization of MOFs, allowing for the introduction of specialized functionalities via a "tag". 102,126 In 2013, Gascon and colleagues conducted a study illustrating that the alteration of NH2-MIL-125(Ti) with dye-like moieties enhances its efficacy for visible light employment. By employing the -NH2 groups in the ligands as markers, dye-like moieties were introduced into the MOF, resulting in the formation of methyl red-MIL-125(Ti) (MR-MIL-125(Ti)). 126 In comparison with the original MOF, MR-MIL-125 displayed a noticeable red shifting, exhibiting an absorption edge extending to nearly 700 nm, which could be owing to the expanded conjugation of the ligand's aromatic system. The enhanced light-harvesting capability of MR-MIL-125(Ti) contributed to its greater performance in the photocatalytic oxidation of benzyl alcohol to benzaldehyde upon exposure to visible light, outperforming the neat MOF.

Garcia and colleagues showed that photogenerated charge carrier separation occurs in NH2-MIL-125 when methyl viologen is used as the e acceptor and N,N,N,N-tetramethyl-pphenylenediamine is used as the e⁻ donor. 102 Their findings confirmed that upon irradiation, NH2-MIL-125(Ti) can concurrently facilitate the redox processes, which are dependent on the redox potential of the probe molecule involved. Such an approach notes a limit: while redox couples like methyl viologen serve a good purpose in characterizing charge dynamics, their electron donor and acceptor will depend on the context and cannot be generalized. Future studies should reflect how photocatalytic efficiencies vary between environments with respect to different electron donor and acceptor combinations of differing redox potential.

Wu and team created an innovative and simple post-grafting method, utilizing a Schiff base chemical interaction amongst the aldehyde and -NH2 groups in MOF to develop aromatic heterocycle-grafted MOF photocatalysts (Fig. 3e).⁵¹ The enhanced conjugation in these MOFs demonstrated reduced E_g s and superior catalytic performance. After modification with aromatic heterocycles, NH2-MIL-125(Ti) exhibited a red shifting in its absorption wavelength. This modification introduced a Schiff base structure formed between the aldehydes and NH₂ groups and on the organic linkers, creating a superconjugation effect with the aromatic benzene ring in the MOF and the aromatic heterocycles. This superconjugation enhances charge excitation and transition during photocatalytic reactions. The post-synthetic incorporation of aromatic heterocycles extended the π -delocalized system, enhancing the separation and transference of photoinduced exciton pairs while preserving the original framework. This led to a marked advancement in visible region absorption efficiency. PSM provides several benefits, such as the incorporation of diverse functional groups, independence from the solvothermal process, easy isolation of modified MOFs, and feasible reaction conditions. Nevertheless, how the changes result in the long-term stability of the MOF for photocatalysis has to be evaluated. The Schiff base structure might degrade under reaction conditions, and improved conjugation may lead to side reactions or changes in the distribution of products. Critical assessment of such PSM concepts has to take place, focusing both on immediate catalytic enhancement and long-term stability in practical applications.

Facet engineering on MILs is another approach for revealing more active metal centers. While this technique is frequently used to improve the surface characteristics of conventional nano photocatalysts, its employment towards soft crystal MOFs is seldom documented. 127-129 Based on recent research, Guo et al. performed the facet engineering of nano-structured NH₂-MIL-125(Ti) by adjusting the amount of the structuredirecting surfactant, CTAB (Fig. 3f). 107 This adjustment increased the occurrence of metal clusters on the material's planes, thereby boosting its photoactivity. Specifically, NH2-MIL-125(Ti) with the active facet (110) exhibited a significantly better quantum yield. It achieved a three-fold better photocatalytic H₂ production rate than that observed with the sample dominated by the (111) facet. Functionalization can be done through various techniques, but maintaining the structural integrity of MOF during modification is important. Unlike traditional nano photocatalysts, which are well known for facet engineering, MOFs create difficulty in their delicate crystal structures and their modifiability. Therefore, the basic challenge lies in the functional enhancements, ensuring that the modifications do not affect the material's stability in practical applications and efficiency.

Photocatalytic application 10.

In recent decades, thanks to its renewable nature, photocatalysis has gained significant attention for applications like H₂ generation via water splitting, CO2 reduction, organic pollutant degradation, and H2O2 production. Photocatalytic reactions occur in three steps (Fig. 4a). 130 (1) Photons with energy exceeding the semiconductor's E_g excite electrons from the VB to the CB, forming e⁻ and h⁺ pairs. (2) These pairs separate and migrate to the surface. (3) The pairs interact with chemical species on the surface, though some may recombine instead of reacting. Ti-centered MOFs, like MIL-125(Ti), are effective for photocatalysis due to the LMCT mechanism, which facilitates charge transfer from the organic ligand to the metal cluster. MIL-125 can only be excited by UV light, while NH₂-MIL-125, made from 2-amino terephthalate, can absorb visible light, improving solar energy utilization. The NH₂ group enhances charge transfer from oxygen to titanium in the TiO₅(OH) cluster, shifting the absorption edge of NH2-MIL-125 to 520 nm, a significant red shift from MIL-125's 350 nm. Table 1 lists the different photocatalytic performances of NH₂-MIL-125 (Ti)-based multifunctional materials.

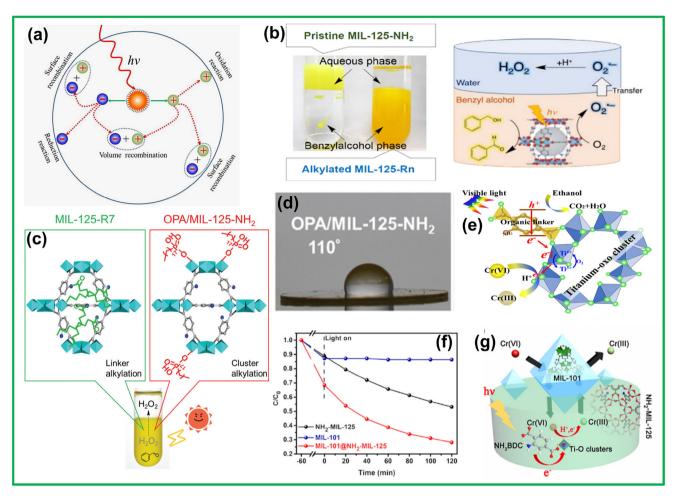


Fig. 4 (a) Photoexcitation and charge decay pathway in photocatalyst. Reprinted with permission from ref. 130. Copyright 2019 Wiley-VCH Verlag GmbH & Co. (b) (left) Digital pictures of two-phase systems composed of aqueous phase and benzyl alcohol phase containing NH2-MIL-125 (left) and MIL-125-Rn (right). (right) Photocatalytic H₂O₂ production utilizing the two-phase system. Reprinted with permission from ref. 131. Copyright 2019 Wiley-VCH Verlag GmbH & Co. (c) Structures of linker-alkylated NH2-MIL-125, MIL-125-R7 (top left), and cluster-alkylated NH2-MIL-125, OPA/ NH₂-MIL-125 (top right). (bottom) The digital picture of the two-phase system composed of a BA/water phase containing OPA/NH₂-MIL-125 in the BA phase. (d) Photographs of water droplets on tablets of OPA/NH2-MIL-125. Reprinted with permission from ref. 132. Copyright 2019 The Royal Society of Chemistry. (e) Schematic illustration of photocatalytic Cr(vi) reduction over NH2-MIL-125(Ti) under visible-light irradiation. Reprinted with permission from ref. 133. Copyright 2014 Elsevier B.V. (f) Adsorption and photocatalytic degradation toward Cr(vi) with MIL-101, NH2-MIL-125, and MIL-101@NH2-MIL-125. (g) Schematic illustration of the enhanced photocatalytic degradation process towards Cr(vi) with micro/mesoporous MIL-101@NH₂-MIL-125 materials. Reprinted with permission from ref. 134. Copyright 2017 Wiley-VCH Verlag GmbH & Co.

10.1. H₂O₂ generation

Hydrogen peroxide (H₂O₂) is a clean oxidant and reductant extensively needed in industrial chemistry and a potential high-energy fuel for fuel cells.¹⁷⁶ H₂O₂ production typically follows two primary routes: a successive two-step single-ereduction ($O_2 \rightarrow O_2^{\bullet-} \rightarrow H_2O_2$) or a one-step two-e reduction $(O_2 \rightarrow H_2O_2)$. In photocatalytic systems, h⁺ in the VB oxidizes H₂O to form O₂ and H⁺ (eqn (1)), while electrons (e⁻) in the CB reduce adsorbed O2 to H2O2. From a photoelectrochemical standpoint, the direct one-step two-e-reduction path has better efficiency for producing H_2O_2 . 177–179

The mechanism of the two-step single-electron pathway involves the reactions discussed in eqn (1)-(5). Alternatively, the one-step, two-electron pathway involves the direct reaction of O_2 with two H^+ to produce H_2O_2 (eqn (6)). The photocatalytic overall response can be summarized as follows (eqn (7)):

$$2H_2O + 4h^+ \rightarrow O_2 + 4H^+$$
 (1)

$$O_2 + e^- \to O^{2^{\bullet -}} \tag{2}$$

$$O^{2^{\bullet-}} + H^+ \to HO_2^{\bullet} \tag{3}$$

$$HO_2$$
 \cdot $+ e^- \rightarrow HO_2^-$ (4)

$$HO_2^- + H^+ \rightarrow H_2O_2 \tag{5}$$

$$O_2 + 2H^+ + 2e^- \to H_2O_2$$
 (6)

$$2H_2O + O_2 \xrightarrow{hv} 2H_2O_2 \quad \Delta G_0 = 117 \text{ kJ mol}^{-1}$$
 (7)

 Table 1
 List of NH₂-MIL-125(Ti)-based multifunctional materials towards different photocatalytic applications

o.	Photocatalytic application	Material	Band gap (eV)	Light source	Photocatalytic activity	R
	H ₂ O ₂	NiO/NH ₂ -MIL-125	_	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	160 μM, 20 h	1
	production	OPA/NH ₂ -MIL-125	_	500 W Xe lamp $(\lambda > 420 \text{ nm})$	_ ` `	1
		MIL-125-R <i>n</i>	~ 2.55	_	_	1
		NH ₂ -MIL-125-D	2.50	200 W metal halide lamp	62.7 μM, 120 min	1
	Cr(vi) reduction	NH ₂ -MIL-125(Ti)	~ 2.65	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	90% reduction, 120 min	1
		M/NH_2 -MIL-125 (M = Pt and Au)	2.2 and 2.38	300 W Xe lamp	_	1
		$MIL-101@NH_2-MIL-125$	2.4	200 W metal halide lamp	88% reduction, 100 min	1
		PPynt@NH ₂ -MIL-125	1.52	300 W Xe lamp ($\lambda > 400 \text{ nm}$)	99.02% reduction, 60 min	1
		NA/NH ₂ -MIL-125(Ti)	2.3	White light (100 W LED)	85% reduction, 90 min	1
)	CO ₂ reduction	NH ₂ -MIL-125(Ti)	2.65	300 W Xe lamp ($\lambda > 400 \text{ nm}$)	HCOO ⁻ = 8.14 mmol, 10 h	1
		Pt/NH ₂ -MIL-125(Ti)	2.2	300 W Xe lamp	HCOOH = 235 μ mol g ⁻¹ h ⁻¹	1
2		Co/NH ₂ -MIL-125(Ti)	2.4	300 W Xe lamp ($\lambda > 400 \text{ nm}$)	$HCOOH = 38.4 \text{ mmol g}^{-1} \text{ h}^{-1}$	1
		Ni/NH ₂ -MIL-125(Ti)	2.5	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	 CO = 12.46 μmol g ⁻¹ h ⁻¹	1
		Bi ₂ S ₃ @NH ₂ -MIL-125	2.1	300 W Xe lamp (λ > 400 nm)		1
,		ZnS/CQDs/NH ₂ -MIL-125(Ti)	2.68	300 W Xe lamp with 420 nm cut-off	CH ₃ OH = 501.57 µmol g ⁻¹ ,	1
		*CO/NH MII 125(Ti)	2.42	filter	CH ₃ CH ₂ OH = 564.70 μ mol g ⁻¹	1
,		rGO/NH ₂ -MIL-125(Ti)	2.43	300 W Xe lamp ($\lambda > 400 \text{ nm}$)	$CH_3OH = 47.2 \text{ mmol } g^{-1} \text{ h}^{-1}$ $CH_3OH = 434.46 \mu\text{mol } g^{-1} \text{ h}^{-1}$,	1
7		Cu ₂ O/Pt/NH ₂ -MIL-125(Ti)	2.68	300 W Xe lamp with 400 nm cut-off filter	$CH_3OH = 434.46 \mu \text{mol g} \cdot \text{I h}$, $CH_3CH_2OH = 718.47 \mu \text{mol g}^{-1} \text{h}^{-1}$	1
3		TiO ₂ @NH ₂ -MIL-125	2.35	—	CO = 159 µmol g ⁻¹	1
,	N ₂ fixation	CH ₃ -MIL-125(Ti)			1.39 µmol g ⁻¹ h ⁻¹	1
	112 HARLION			off filter		
)		OH-MIL-125(Ti)	_	300 W xenon lamp with 400 nm cut- off filter		1
1		NH ₂ -MIL-125(Ti)	_	300 W xenon lamp with 400 nm cut- off filter	12.25 μ mol g ⁻¹ h ⁻¹	1
2		Ce-NH ₂ -MIL-125	2.66	300 W xenon lamp (light intensity of 200 mW cm ⁻²)	$39.4 \ \mu mol \ g^{-1} \ h^{-1}$	1
3	H ₂ evolution	PtNPs@NH ₂ -MIL-125(Ti)	_	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	367 μmol g ⁻¹ h ⁻¹	
	112 0101411011	Pt/NH ₂ -MIL-125/(OH) ₂	~2.2	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	707 µmol g ⁻¹ h ⁻¹	
		NH ₂ -MIL-125/OH	_	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	377 umol g ⁻¹ h ⁻¹	
		NH_2 -MIL-125/(NH_2) ₂	~1.6 eV	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	253 μmol g ⁻¹ h ⁻¹	
		CdIn ₂ S ₄ @NH ₂ -MIL-125	_	300 W xenon lamp ($\lambda > 400$ nm)	2550 μmol g ⁻¹ h ⁻¹	
		CdS/NH ₂ -MIL-125(Ti)	_	350 W xenon lamp ($\lambda > 420 \text{ nm}$)	6.62 mmol g ⁻¹ h ⁻¹	
		Pd/NH ₂ -MIL-125(Ti)/CdIn ₂ S ₄	2.08	300 W xenon lamp $(\lambda > 420 \text{ nm})$	$4.60 \text{ mmol } g^{-1} \text{ h}^{-1}$	
		ZnIn ₂ S ₄ @NH ₂ -MIL-125(Ti)	2.48	300 W xenon lamp $(\lambda > 420 \text{ nm})$	2204.2 μmol g ⁻¹ h ⁻¹	
		NH ₂ -MIL-125(Ti)@ZnIn ₂ S ₄ /CdS	1.84	300 W xenon lamp $(\lambda > 400 \text{ nm})$	2.367 mmol g ⁻¹ h ⁻¹	-
		NH ₂ -MIL-125(Ti)/Ti ₃ C ₂ QDs/ ZnIn ₂ S ₄	2.35	Visible light	2931.9 μmol g ⁻¹ h ⁻¹	-
		1T-MoS ₂ /NH ₂ -MIL-125	_	Xenon lamp (λ ≥ 420 nm)	1454 μmol g ⁻¹ h ⁻¹	-
Į.		Ni ₂ P/NiO@NH ₂ -MIL-125(Ti)	_	Visible light	1611 μmol g ⁻¹ h ⁻¹	-
		Co@NH ₂ -MIL-125(Ti)	2.59	500 W Hg/Xe lamp ($\lambda > 408 \text{ nm}$)	_	
		NH ₂ -MIL-125@TiO ₂	_	300 W xenon lamp ($\lambda \ge 420 \text{ nm}$)	496 μmol g ⁻¹ h ⁻¹	
		NH ₂ -MIL-125(Ti)/Pt/g-C ₃ N ₄	2.6	300 W Xe lamp with 380 nm UV cutoff filter	3986 μmol g ⁻¹ h ⁻¹	
	Drug	Oxygen-enriched vacancy NH ₂ - MIL-125(Ti)	2.36	300 W Xe lamp	TCH = 80%, 2 h	
	degradation	Bi ₃ O ₄ Br/NH ₂ -MIL-125(Ti)	2.8	300 W xenon lamp ($\lambda > 400 \text{ nm}$)	TCH = 88.5%, 90 min	
		g-C ₃ N ₄ /NH ₂ -MIL-125	2.56	UV LED irradiation (384 nm)	DCF = 100%, 2 hours	
		LaFeO ₃ /NH ₂ -MIL-125	2.4	150 W xenon lamp	CAF = 87% and CBZ = 74%, 1 hour	
		NH ₂ -MIL-125@CoFe PBA	2.59	300 W Xe lamp ($\lambda \ge 420 \text{ nm}$)	TCH = 86.9%, 2 hours	
		UiO-66/NH ₂ -MIL-125/g-C ₃ N ₄	2.68	Visible light (300 W xenon lamp, λ > 420 nm)	_	
		NH ₂ -MIL-125@MIL-88B	_	300 W Xe lamp	TCH = 96.26%, 120 minutes	
		Ag/NH ₂ -MIL-125@MIL-88B	2.35	Simulated sunlight (350 W Xe lamp)	KP = 94.2%, 180 minutes	
		Pt/NH ₂ -MIL-125(Ti)	2.48	Simulated solar light	— 54.270, 100 minutes	
		Pd/NH ₂ -MIL-125(Ti)	2.50	Simulated solar light	_	
		Ag/NH ₂ -MIL-125(Ti)	2.70	Simulated solar light	_	
	Dye disposal	Au@NH ₂ -MIL-125(Ti)/CdS	2.28	300 W xenon lamp, $\lambda \ge 420 \text{ nm}$	MB = 93.3%, 30 min	
	J I	NH ₂ -MIL-125(Ti)/CdS/Graphene	2.2	300 W xenon lamp ($\lambda > 420 \text{ nm}$)	RhB = 95%, 30 min	
		NH ₂ -MIL-125(Ti) nanodots on CF/ MoS ₂	_	300 W xenon lamp	AO7 = 67.9%, MB = 94.3%, 120 MIN	
		BiOI/ZnFe ₂ O ₄ /g-C ₃ N ₄	_	LED light	AB92 = 80%, MB = 73.8%, phenol: 66.6%,	
		$\mathrm{NH_2} ext{-}\mathrm{MIL} ext{-}125$ @BiVO $_4$	_	300 W xenon lamp ($\lambda > 420 \text{ nm}$)	120 min MG = 93.6%, MB = 97.6%, CV = 97.9%,	
		${\rm Ag_3PO_4@NH_2\text{-}MIL\text{-}125}$	2.39	Philips 100 W equivalent daylight	CR = 58.2%, 3 hours MB = 93.3%, RhB = 95.5%, 60 min	
		Ag-VO.@NHMII-125	2.27	lamp UV and visible light	MB = 99.68% (MB), RhB = 102.38%, 60 min	
		$Ag_3VO_4@NH_2$ -MIL-125 $Ag_2WO_4@NH_2$ -MIL-125	2.27	UV and visible light	MB = 49.12%, RhB = 38.55%, 60 min	
		Ag/AgBr/NH ₂ -MIL-125 Ag/AgBr/NH ₂ -MIL-125(Ti)	2.6	300 W xenon lamp ($\lambda > 420 \text{ nm}$)	MB – 49.12%, KHB – 38.55%, 60 HHH MO = 70%, 120 min	
		CQDs/P25/NH ₂ -MIL-125	2.0	300 W xenon lamp with an AM1.5 G	MB = 90.4%, 150 min	
		-		optical filter 300 W xenon lamp ($\lambda \ge 420 \text{ nm}$)	RhB = 100%, 45 min	

The generation of H₂O₂ from H₂O and O₂ via photocatalysis is an energetically demanding reaction characterized by a standard ΔG_0 of 117 kJ mol⁻¹.

NH2-MIL-125 was produced and subsequently modified with NiO to create NiO/NH2-MIL-125 via the method outlined by Isaka et al. 135 The UV-DRS spectrum shows high-energy absorption due to $\pi \to \pi^*$ transitions of the organic ligands and a visible region absorption from LMCT involving NH2groups on the ligands, confirming that both NH2-MIL-125 and NiO/NH₂-MIL-125 are responsive to visible light for photoactive reactions. Photocatalytic H₂O₂ generation, using NH₂-MIL-125 in oxygen-saturated acetonitrile with TEOA as an electron donor under visible light, showed no H2O2 without light or TEOA. The H₂O₂ production rate improved significantly with NiO/NH₂-MIL-125 or Pt/NH₂-MIL-125 compared with NH₂-MIL-125 alone. Oxidation of TEOA produced aldehyde and its hydrated form. Stoichiometric analysis of H₂O₂ and electron donor association was complicated by multiple alcohol moieties in TEOA, so photocatalytic H₂O₂ production was carried out by oxygen reduction and oxidation of benzyl alcohol. NH2-MIL-125 resulted in H₂O₂ being only 5.7% of the benzaldehyde generated after 8 hours, while NiO/NH2-MIL-125 increased this yield to 93%. NiO/NH2-MIL-125 also showed half the H2O2 decomposition rate of NH2-MIL-125 at high H2O2 concentrations, suggesting repression of NiO's two-electron reduction as the cause for the enhanced selectivity. EPR experiments with DMPO detected O2 - species, confirming their formation under visible light with NH2-MIL-125 but with a reduced rate for NiO/NH2-MIL-125. These results show that NiO accelerates the conversion of O₂ to H₂O₂. The higher H₂O₂ selectivity in NiO/NH₂-MIL-125 is attributed to the faster decomposition of O2, while under visible light, photogenerated electrons reduce Ti(IV) to Ti(III), forming Ti₈O₈(OH)₄, which reduces O₂ to superoxide that forms H₂O₂ via disproportionation, further enhanced by NiO. Although NiO/NH2-MIL-125 efficiently catalyzed H₂O₂ production in a single-phase system, its performance declined during recycling tests due to the instability of NH₂-MIL-125, a typical issue with Ti-based MOFs that emphasizes the requirement for robust materials or surface modifications that can prevent the degradation of the framework without a reduction in photocatalytic efficiency. Furthermore, using a single-phase system for both H₂O₂ production and catalyst mixing complicates H₂O₂ collection and leads to its catalytic decomposition, requiring an additional separation process to isolate H₂O₂ from the MOFs.

To address the mixing issues of H2O2 and MOFs, a twophase system of benzyl alcohol (BA) and H2O was considered for photocatalytic H₂O₂ production (Fig. 4b). ¹³¹ This system separates MOFs in the BA phase and H2O2 in the aqueous phase via NH2-MIL-125's hydrophobicity. PSM added alkyl chains to the MOF, creating MIL-125-Rn (n = 4 and 7). Alkylation caused a blue shift in UV-vis DRS absorbance but maintained visible region absorption. NH2-MIL-125 dispersed into the water phase, with H2O2 generation only in the aqueous phase. MIL-125-Rn (n = 4 and 7) dispersed into the BA phase, with H_2O_2 formation in the aqueous phase. MIL-125-Rn

showed significantly higher H₂O₂ production than NH₂-MIL-125, while benzyl alcohol oxidation rates were similar, indicating selective H2O2 production increased activity. In a single-phase system, H₂O₂ decomposes into OH⁻ and OH, suppressing its production, but this issue is avoided in the biphase system. 132 Seawater further enhances H₂O₂ production by stabilizing O2 - through Na complexation. Under visible light, MIL-125-Rn generates electrons that reduce Ti(IV) to Ti(III) via LCCT, forming Ti₈O₈(OH)₄-, which reduces O₂ to superoxide, then forms H₂O₂ through disproportionation, accelerated by H⁺ or Na⁺. MIL-125-R7's H₂O₂ production is four times greater than NH2-MIL-125 but still limited, likely due to a reduced surface area (560.7 m² g⁻¹ vs. 1498 m² g⁻¹ for NH₂-MIL-125), resulting from alkylation (Fig. 4c). Inclusion of alkyl chains in MOFs can, however, result in trade-offs between their hydrophobicity and surface area, thus reducing the available sites for the adsorption of reagents and for catalytic activity. This reveals a significant lacuna in balancing hydrophobic and surface area parameters for realizing high photocatalytic efficiencies while limiting decomposition rates. To counteract the reduction in SSA of the parent MOF due to the alkylation of the linker, an additional hydrophobic MOF, OPA/NH2-MIL-125, was developed by alkylating the Ti cluster with octadecyl phosphonic acid (OPA). In MIL-125-R7, the alkyl chains filled its pores, while in OPA/NH2-MIL-125, OPA modified the Ti₈O₈(OH)₄ cluster only on the exterior part, preserving most pores. This selective modification maintained higher photocatalytic activity for H₂O₂ production, attributed to the rapid reactant's diffusion and product generation. The specific surface area of OPA/NH2-MIL-125 was measured at 1242 m² g⁻¹, showing that the majority of the SA was preserved after metal cluster alkylation, with the authentic pore diameter and volume remaining intact. OPA/NH2-MIL-125's water contact angle, 110°, indicates its hydrophobic nature (Fig. 4d). In a benzyl alcohol/H₂O bi-phasic system, the OPA/NH₂-MIL-125 selectively dispersed in the benzyl alcohol phase. At the same time, H₂O₂ generation took place in the water phase with visible-light radiation. H₂O₂ generation with OPA/NH₂-MIL-125 was threefold greater in comparison with MIL-125-R7 after three hours of irradiation. Additionally, the photocatalytic activity of OPA/NH₂-MIL-125 in the benzaldehyde phase was 1.7 times more in comparison with MIL-125-R7 as they diffused swiftly through unblocked pores. These findings prove that cluster-alkylated OPA/NH2-MIL-125 has better efficiency for photocatalytic H2O2 generation.

Zhang et al. reported the formation of NH2-MIL-125(Ti) and NH₂-UiO-66(Ti) isomers via simple hydrothermal and post-synthesis exchange (PSE) methods. 136 Upon high-temperature treatment, defective NH2-MIL-125(Ti) (NH2-MIL-125-D) was formed, featuring a relatively short Ti-O bond length advantageous for charge transportation, thus improving photocatalytic performance compared with NH2-UiO-66(Ti). NH2-MIL-125-D has extra active sites owing to the loss of ligands, creating defective titanium sites that enhanced the reduction of adsorbed O₂ to H₂O₂ and increased yield. The N₂-adsorption isotherms showed higher adsorption for NH₂-MIL-125(Ti) than

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NH₂-UiO-66(Ti), attributed to topological properties, allowing more significant oxygen adsorption. In contrast, NH₂-MIL-125-D showed lower N_2 uptake with hysteresis around 0.4-0.8 P/P_0 , indicating mesopore formation due to high-temperatureinduced ligand loss. Photocatalytic H₂O₂ production in water with catalysts (15 mg in 50 ml) under visible light showed slightly enhanced activity for NH2-MIL-125(Ti) over NH2-UiO-66(Ti), while NH₂-MIL-125-D significantly enhanced H₂O₂ production, achieving 62.7 mM H₂O₂ after 120 min. The increased performance of OPA/NH2-MIL-125 can be ascribed to the presence of additional active sites created by defect-rich titanium resulting from ligand shortage in the framework. These defects facilitated the adsorption of oxygen and its rapid reduction to H2O2. However, there are critical issues in maintaining structural integrity during repeated use and in the effective separation of H2O2 from the catalyst to avoid decomposition. Furthermore, improvements in H₂O₂ collection techniques and the development of a biphasic system with hydrophobic treatments spatially separate the catalyst and H₂O₂, thus enhancing the efficiency and selectivity.

10.2. Cr(vi) reduction

The Cr(v1) photoreduction typically involves photogenerated electrons (e⁻) in the conduction band (CB) of a semiconductor, with the CB edge needing a more negative potential than the Cr(vi)/Cr(III) redox potential. Direct reduction to Cr(0) is not possible due to the more negative Cr(III)/Cr(0) potential. e can reduce Cr(vi) either directly or via intermediates like Cr(v) and Cr(IV), following two main pathways: a one-step, three-electron transfer to Cr(III) (eqn (8)) or a stepwise, single-electron transfer (eqn (9)). 180-182 Acidic conditions are favored, and hole scavengers are added to prevent reoxidation of Cr(III) by ROS, improving e separation and reaction efficiency. Without scavengers, the reaction is slower in pure water due to recombination, leading to a four-electron transfer process (eqn (10)). 183, 184

$$Cr_2O_7^{2-} + 14H^+ + 6e^- \rightarrow 2Cr^{3+} + 7H_2O$$
 (8)

$$Cr(vi) + e^- \rightarrow Cr(iv) + e^- \rightarrow Cr(iii)$$
 (9)

$$Cr(III) + ROS \rightarrow Cr(VI)$$
 (10)

Wang and colleagues conducted research on the Cr(vi) photoreduction under visible light with NH2-MIL-125(Ti), which was synthesized using an adapted solvothermal method in comparison with MIL-125(Ti). MIL-125(Ti) is a well-crystallized white powder with particle sizes around 2 nm, while NH₂-MIL-125(Ti) is a yellow powder with particles ranging from 400 to 600 nm. XRD analysis of NH2-MIL-125(Ti) showed no peaks related to bulk TiO2 anatase or rutile phases, indicating the disappearance of large titanium oxide aggregates. UVvis DRS spectroscopy revealed that MIL-125 absorbs light up to 350 nm, while NH₂-MIL-125 extends absorption to about 520 nm due to the 2-amino terephthalate ligand, enhancing its visible light absorption. Photocatalytic optimization studies showed that at pH 2.1, adding the hole scavenger ethanol

increased the photocatalytic Cr(vI) reduction efficiency of NH₂-MIL-125(Ti) from 76% to 91%. In contrast, citric acid and EDTA reduced the efficiency to 72% and 66%, respectively. Ethanol, absorbed into NH2-MIL-125(Ti), aids its mineralization and enhances Cr(vI) reduction by capturing photogenerated holes (egn (11)-(13)), improving electron utilization. Conversely, citric acid and EDTA, through hydrogen bonding with NH₂-MIL-125(Ti), promote electrostatic adsorption of Cr(vi) but limit electron transfer at Ti sites, reducing the Cr(vi) reduction efficiency.

$$NH_2$$
-MIL-125(Ti) + $h\nu \rightarrow NH_2$ -MIL-125(Ti)($h^+ + e^-$) (11)

$$CH_3CH_2OH + h^+ \rightarrow CH_3CH_2OH^+ \rightarrow CO_2 + H_2O + others$$

$$(12)$$

$$14H^{+} + Cr_{2}O_{7}^{2-} + 6e^{-} \rightarrow 2Cr^{3+} + 7H_{2}O$$
 (13)

The Cr(vi) reduction efficiency by NH₂-MIL-125 lessened significantly as the pH increased from 4.07 to 8.17. Under acidic conditions, the photocatalytic reaction follows eqn (13), where the abundance of H⁺ ions enables the reduction of Cr(v_I). Conversely, in basic environments, eqn (14) predominates as CrO₄²⁻ becomes the dominant species. The zeta potential values of both MIL-125 and NH2-MIL-125 become increasingly negative along with rising pH levels. NH2-MIL-125(Ti) exhibits a positive zeta potential below pH 4.6, which enhances the adsorption of anionic Cr(vi) and consequently improves photoreduction efficiency. At higher pH values, the catalyst surface becomes negatively charged, repelling the anionic Cr(vi) entity and promoting the cationic Cr(III) adsorption entity. Moreover, at pH > 6, there is precipitation of Cr(III) as $Cr(OH)_3$, which can mask active sites on NH2-MIL-125.

$$CrO_4^{2-} + 4H_2O + 3e^- \rightarrow Cr(OH)_3 + 5OH^-$$
 (14)

XPS analysis confirmed the valence states of chromium bound to the MOF. Peaks observed at 577.3 eV and 580.7 eV for Cr 2p_{3/2} orbitals indicate the presence of Cr(III), suggesting the reduction of Cr(vI) to Cr(III). Additionally, a peak at 586.7 eV corresponds to Cr(vi) absorbed onto the titanium MOF, attributed to the material's large specific surface area, pore assembly, and electrostatic interactions. This was further supported by a shift in the N 1s peak in Cr/NH₂-MIL-125(Ti). The projected mechanism (Fig. 4e) comprises the enhancement of light harvesting by H₂ATA in NH₂-MIL-125. Photoexcited H₂ATA transfers photoinduced e⁻s to the Ti-O clusters, producing Ti³⁺. These Ti³⁺ ions re-oxidize to Ti⁴⁺ upon exposure to air. The charge transfer between Ti3+ and Ti⁴⁺ in the Ti-O clusters was clarified by Horiuchi et al. using ESR and UV-Vis spectra.⁵⁶ Ethanol, which was adsorbed onto the NH₂-MIL-125, functions as a hole scavenger, decomposing into CO2 and H2O, thereby creating a basic environment that promotes the Cr(v1) reduction. NH2-MIL-125 also demonstrated good reusability, with only an 8% reduction in photoactivity after five consecutive cycles of washing with HNO₃, deionized water, and drying. Furthermore, ethanol enhanced Cr(vi) reduction efficiency from 76% to 91%, while the exact mechaNanoscale

nism of stabilizing the separation of electron and hole remains to be studied. Decoding the pH dependency in Cr(v1) reduction is also important, as its higher efficiency under acidic conditions reduces catalyst reusability, owing to Cr(III) precipitation. Thus, pH-stabilized photocatalysts or dual-function photocatalysts could be investigated to maintain high activity over a wider pH range, thereby minimizing Cr(III)

precipitation. Modification of MIL-125 and NH2-MIL-125 with highly dispersed platinum nanoparticles (3-9 nm) resulted in improved Cr(vi) photoreduction, as the Pt nanoparticles accelerated electron transfer and delayed electron-hole recombination. 137 In 2017, Kitagawa et al. developed a MOF-on-MOF hybrid heterostructure using the internal extended growth method (IEGM). 134 Typically, epitaxial growth requires MOFs with similar crystallographic parameters. However, IEGM overcomes lattice mismatching limitations, enabling the synthesis of MIL-101(Cr) over NH2-MIL-125(Ti). This process involves microwave treatment of pre-made MIL-101(Cr) and NH₂-MIL-125(Ti) precursors with PVP as a structure-directing entity, resulting in MIL-101(Cr) nanoparticles homogeneously distributed on NH2-MIL-125 nano tablets. The synthesized micromesoporous MOF-on-MOF system, MIL-101@NH2-MIL-125, demonstrated a substantially higher Cr(vI) adsorption capacity (3.16 mg g⁻¹) in comparison with NH₂-MIL-125(Ti) and MIL-101(Cr). After being exposed under visible light for two h, MIL-101@NH2-MIL-125 accomplished a 72% reduction in Cr(vI), whereas NH₂-MIL-125(Ti) achieved a reduction of 47%, and MIL-101(Cr) exhibited negligible reduction (Fig. 4f and g). This improved photocatalytic performance is credited to the harmonious interface among NH₂-MIL-125 and MIL-101(Cr), with the mesopores in MIL-101(Cr) significantly enhancing Cr(vi) adsorption.

Defect engineering in photocatalysts is a promising approach aimed at enhancing the number of active sites and promoting effective photogenerated exciton separation. 185-187 In 2021, researchers developed a defective NH2-MIL-125(Ti) with linker vacancies over controlled thermal treatment, which involved decomposing the carboxylic acid group of 2-amino terephthalic acid. 188 Through prolonged thermal treatment, the decomposition of the NH2-MIL-125 ligand progressed inward from the surface, leading to an increased presence of linker-vacancy defects. This modification significantly impacted the material's framework structure, light-harvesting properties, and energy band structure. Moderate levels of linker-vacancy defects have been found to enhance the photoinduced charge carrier separation significantly. This enhancement resulted in a Cr(v1) photoreduction that was 9.2 times higher compared with the neat MOFs. However, high concentrations of linker-vacancy defects are known to deteriorate the crystal structure and impede efficient charge separation. There should be a balance between too much and too little defect content, as too many will destabilize the framework, whereas too few will not effectively improve photocatalytic activity. This balance is sensitive and needs further study.

Conductive polymers such as polyaniline and polypyrrole have garnered considerable interest owing to their versatile applications, including their ability to reversibly switch between oxidation and reduced states, which is advantageous for reducing toxic metal ions such as hexavalent chromium. 189 Choe et al. established a core-shell heterostructure called polypyrrole nanotube@NH2-MIL-125 (PPynt@NH2-MIL-125) by growing NH₂-MIL-125 in situ on polypyrrole nanotubes. ¹³⁸ The PPynt@NH2-MIL-125 heterostructure exhibited outstanding outcomes in Cr(v1) photoreduction under visible light conditions. Incorporating 5 wt% of PPvnt resulted in a superior elimination efficiency, approximately 70% improvement compared with pure NH2-MIL-125. Moreover, these nanocomposites demonstrated excellent robustness and cycling stability. The study puts the critical role played by MOFs in the spotlight in enhancing the photocatalytic performance of NH₂-MIL-125 for Cr(vi) photoreduction, specifically by facilitating the effective charge carrier separation under visible light in the PPvnt@NH2-MIL-125 nanocomposite.

In 2023, Wang et al. fabricated a NA/NH2-MIL-125(Ti) homojunction by partially post-synthesis modifying the Ti-MOF with 1-naphthylamine (-NA, a chromophoric group). 139 Unlike the earlier reported MR/NH2-MIL-125(Ti), the conjugated N=N bond and the naphthalene ring offered advanced photoactivity. Photoreduction of Cr(vI) using white light showed NA/NH2-MIL-125 had the highest activity among tested MOFs, with dark control experiments confirming the necessity of photoexcitation. Physical mixture experiments suggested a tightly bound complex, supporting its superior performance. Pseudo-first-order kinetics analysis revealed MIL-125 had a rate constant 1.5 times greater than NH₂-MIL-125. The NA-group was designed via mixing 1-naphthylamine with diazonium salt, with an optimal NaNO2: amino group ratio of 0.1:1. Lower pH favored Cr(v1) reduction, and EDTA-2Na performed like a hole scavenger and influenced reaction efficiency. NA/NH2-MIL-125 exhibited wide-spectrum responsiveness, with peak performance under purple light (410-420 nm). Incorporating the -NA chromophore notably improved light absorption, charge separation, and the formation of active species. The homojunction with optimal ratio demonstrated the highest photocurrent, and the synthesized samples showed excellent reusability over multiple operation cycles.

Tin disulfide (SnS₂) is a favorable semiconductor owing to its appropriate E_g in the visible region, strong light absorption efficiency, abundance, and resistance to photocorrosion, unlike CdS semiconductors.^{151,190–192} Sun *et al.* demonstrated that SnS₂ nanosheets loaded onto NH₂-MIL-125 form a new Z-scheme heterostructure that effectively reduces $Cr(v_I)$ in visible light.¹⁹³ The photocatalytic efficiency of NTS₂ for $Cr(v_I)$ reduction is influenced by pH, with acidic conditions enhancing reduction efficiency due to increased $Cr(v_I)$ adsorption and electron capture. In contrast, alkaline conditions lead to $Cr(OH)_3$ precipitation, covering active sites and reducing activity. Metal ion leaching, such as Sn from NTS₂, increases with pH due to sodium stannate formation in alkaline environ-

ments. However, Cr(vi) reduction is of utmost effectiveness in strongly acidic media, minimizing Sn ion leaching. NTS₂ environ exhibited excellent recyclability, maintaining a high elimin-

exhibited excellent recyclability, maintaining a high elimination rate of 90.28% after four cycles, with XRD patterns indicating structural stability. The Z-scheme heterojunction facilitated electron transfer from SnS₂ to NH₂-MIL-125, enhancing electron-hole pair separation. Photon absorption by NT and SnS₂ generated electron-hole pairs, followed by electron transfer and Cr(v1) reduction to Cr(III). However, challenges include inefficiency observed at higher pH levels due to electrostatic repulsion and the precipitation of Cr(OH)3, which masks the active sites. Much remains to be discovered about intermediate species, such as Cr(v) and Cr(v), which are generated during the stepwise reduction of Cr(v_I). The combined application of in situ real-time tracking and computational modeling could provide valuable insights into designing photocatalysts that selectively reduce Cr(v1) with the fewest possible side reactions. Optimizing conditions, such as using scavengers and maintaining acidity, alongside exploring hybrid structures, defect engineering, and core-shell heterostructures, can enhance electron transfer, charge separation, and stability.

10.3. CO₂ reduction

The challenge of converting carbon dioxide (CO₂) into valuable chemicals is both formidable and essential for advancing sustainable energy and addressing climate change. CO2's inherent stability, combined with significant thermodynamic and kinetic barriers, makes its transformation particularly difficult. Having a Gibbs free energy change (ΔG°) of -400 kJ mol^{-1} , CO₂ requires substantial energy input and effective catalysis to undergo efficient conversion [C2]. It has a linear molecular structure characterized by two carbon-oxygen double bonds. This inherent stability poses significant challenges to its transformation into useful chemicals without the assistance of a catalyst and the input of energy. 194 Both thermodynamic and kinetic constraints, combined with a substantial energy gap between the LUMO and HOMO of 13.7 eV, besides a superior electron affinity of -0.6 ± 0.2 eV, contribute to the inert nature of CO₂. 195 Photocatalytic CO₂ reduction pathways vary based on the reductant, with carbon-free reductants (H₂O and H₂) generally forming C1 products and carbon-rich entities (CH₄ and CH₃OH) potentially forming C2 and C3 products. Common CO₂ photoreduction products include CO, HCOOH, HCHO, CH₃OH, CH₄, and C₂H₅OH. ¹⁹⁶⁻¹⁹⁹ Scheme 1 illustrates the required reducing electrons, oxidizing holes, and protons (H⁺) for forming these main products (Fig. 5a).²⁰⁰

A promising approach for CO₂ reduction involves creating materials that work under visible light. Photoactive titanium-based MOFs, like NH₂-MIL-125, are effective. Li *et al.*⁵⁰ showed that NH₂-MIL-125, as a semiconductor, converts CO₂ to formate using triethanolamine (sacrificial agent). Amino groups in the framework created an LMCT band that absorbed visible light, generating a long-lived exciton pair. Electrons moved to Ti⁴⁺ ions, reducing them to Ti³⁺, which acted as active centers for CO₂ reduction, while TEOA served both as an e⁻ donor and Lewis base. The team also studied NH₂-MIL-125

in combination with noble metals like Pt and Au in a saturated environment of CO₂. Noble metal NPs embedded in the MOF acted as electron traps, enhancing redox reactions and, thus, overpotential reduction. Over 8 hours, Pt/NH₂-MIL-125 showed a 21% increase in CO₂ reduction compared with unmodified NH2-MIL-125, while Au/NH₂-MIL-125 had a 16% decreased performance with a reduction yield of 9.06 µmol. The stronger Pt-O bond stabilized Pt, facilitating Ti³⁺ formation, whereas the weaker Au-O bond led Au to prefer NH₂ sites, preventing this effect. Selecting the right MOF-metal pair is crucial for boosting photocatalysis. Such difference displays how sensitive the photocatalytic system is to the metal-support interactions, thereby emphasizing the importance of the appropriate choice of metal nanoparticles.

Non-noble metal doping offers distinct advantages over noble metal doping in NH2-MIL-125 for photocatalytic CO2 reduction owing to the abundance and cost-effectiveness of non-noble metals, making large-scale production more economically viable. In 2017, Fu et al. introduced Co-doped NH₂-MIL-125 photocatalysts for visible-light-assisted reduction. 141 The embodiment of Co nanoparticles (NPs) into the MOF framework enabled these particles to act as efficient electron traps, capturing electrons from the excited ligands. Furthermore, replacing triethanolamine with benzylic alcohols as electron donors significantly enhanced the reduction process. Co/NH₂-MIL-125(Ti) demonstrated dual catalytic activity: it facilitated the reduction of CO2 to formic acid (HCOOH) while selectively oxidizing benzylic alcohols to their corresponding aldehydes. This dual functionality enhances the economic and environmental viability of the process. To address the challenge of regulating electron transfer within metal-oxo clusters of MOFs via heteroatom doping, Gao et al. selected Ni for its high electron affinity. They investigated the impact of Ni on electron distribution in NH2-MIL-125 by preparing NH2-MIL-125-Nix/Ti materials with precise Ni concentrations through in situ doping. 142 Nickel, being more electronegative than Ti, alters the electronic arrangement of Ti-O clusters in NH2-MIL-125. PDOS calculations (Fig. 5b) showed a reduced organic ligand contribution in Ni²⁺-doped NH₂-MIL-125's HOMO and LUMO, with increased Ni 3d and Ti 3d contributions. Ni2+ created states between HOMO and LUMO, facilitating electron transfer and enhancing charge separation. Doping with Ni²⁺ significantly enhanced CO₂ to CH₄ conversion. The addition of Ni2+ dopants significantly enhanced the photocatalytic performance, resulting in a three-fold increase in CO yield with 0.5% Ni²⁺ doping and a six-fold increase with 1% Ni²⁺ doping in NH₂-MIL-125 compared with the undoped counterpart. NH2-MIL-125-Ni1%/Ti showed 98.6% CO selectivity due to enhanced charge carrier separation and conductivity. EPR spectra indicated increased exciton separation, more photoinduced holes, and Ti3+ signals, which enhanced electron production and charge separation, boosting photoelectric conversion efficiency and photocatalytic performance.

ZnS is an effective, low-cost semiconductor photocatalyst with a wide $E_{\rm g}$. However, reverse electron flow and recombination can hinder its performance. Combining ZnS and NH₂-

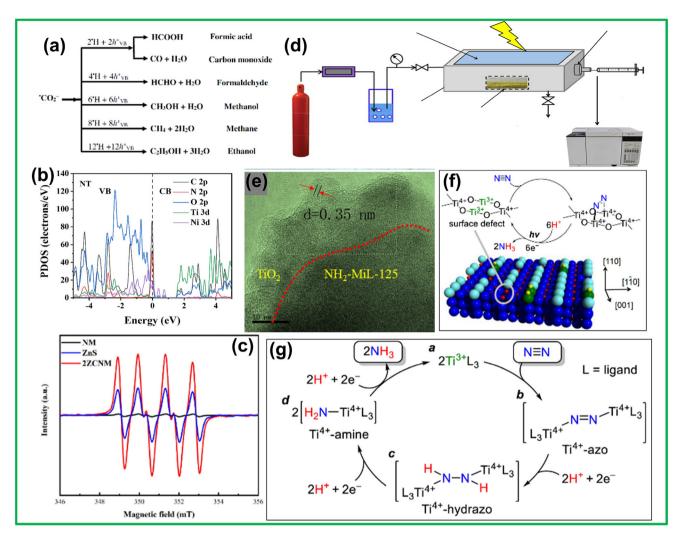


Fig. 5 (a) The total pathways for production of the main products of photocatalysis of CO₂. Reprinted with permission from ref. 200. Copyright 2016 Elsevier Ltd. (b) PDOS for Ni²⁺-doped Ti-MOF indicated as NT. Reprinted with permission from ref. 142. Copyright 2020 Elsevier B.V. (c) ESR spectra by adding DMPO to capture 'O2" with NM, ZnS, and 2ZCNM. Reprinted with permission from ref. 144. Copyright 2023 Elsevier B.V. (d) A schematic illustration of the photoreactor set-up for photocatalytic CO₂ reduction in a gas phase system. Reprinted with permission from ref. 201. Copyright 2021 Elsevier Ltd. (e) HR-TEM image of CNMT-40. Reprinted with permission from ref. 147. Copyright 2023 Elsevier B.V. (f) Proposed photocatalytic cycle for N_2 fixation on the rutile TiO_2 (110) surface. (g) Catalytic cycle for N_2 fixation by Ti^{3+} -containing complexes. Reprinted with permission from ref. 202. Copyright 2017 American Chemical Society.

MIL-125 in a heterojunction addresses these issues and improves performance.206 Mediating agents can further boost photocatalytic activity and carrier separation. 207,208 In 2024, Zhou et al. used carbon quantum dots (CODs) derived from waste biomass sludge as mediators to link ZnS and NH2-MIL-125 towards CO2 photoreduction.¹⁴⁴ This Z-scheme approach treats waste muck and enhances charge transfer, offering two-fold benefits. 209,210 The 2ZCNM composite catalyst showed significant photocatalytic performance enhancements for CO2 reduction, yielding 501.57 μ mol g⁻¹ of CH₃OH and 564.70 μ mol g⁻¹ of CH₃CH₂OH in 3 hours. The improvement was due to the synergistic effects of CQDs and ZnS, which altered the pore structure and introduced mesoporous features, enhancing H2O and CO2 adsorption. While ZnS boosted photocatalytic activity, excess loading blocked active sites, causing electron-hole recombination. The formation of CH₃OH and CH₃CH₂OH was exclusively from CO₂, confirming process selectivity and efficiency. The 2ZCNM composite showed minimal activity loss after five cycles, highlighting its stability as a CO₂ reduction photocatalyst. ESR spectroscopy (Fig. 5c) confirmed the Z-scheme of 2ZCNM, enhancing carrier separation and O2 - radical generation. NM and ZnS created exciton pairs with visible light, while CQDs facilitated electron transfer from NM's LUMO to ZnS's CB. Electrons in ZnS's CB reacted with CO2 to form CH3OH and CH3CH2OH, and NM's HOMO generated hydroxyl radicals (*OH), improving charge separation and CO2 reduction performance. While harnessing CQDs has shown promise in improving carrier separation and enhancing the stability of photocatalysts, optimizing the balance between material loading and undue recombination still presents challenges.

Publications on graphene-assisted MOF composites for visible light photocatalysis are limited. However, reduced graphene oxide (rGO) has exceptional properties such as high thermal conductivity, 211 chemical robustness, 212 large surface area, 213 mechanical strength, 214 charge carrier mobility, 215 and optical transparency, 216,217 making it ideal for composites with NH₂-MIL-125. Olowoyo et al. investigated the photoreduction efficacy of pristine MOF and RGO-NH2-MIL-125 composites for CO2 reduction in a CH3CN + H2O medium with TEOA as a sacrificial reagent in 2020.145 They found that RGO-NH2-MIL-125 showed only marginal increases in methane and ethylene production compared with NH2-MIL-125 but significantly enhanced methanol selectivity in both liquid and gas phases. This performance was due to RGO-NH2-MIL-125's greater visible light absorption and narrower E_g , leading to more efficient solar spectrum use. RGO also improved charge transport and separation, boosting product formation. Computational studies confirmed charge localization and CO₂-TEOA interactions with the MOF, supporting RGO-NH₂-MIL-125's potential as a potent CO2 reduction photocatalyst under visible light exposure. The alignment of CB and VB plays a pivotal part in facilitating charge carrier transfer within heterostructures. The band positions of g-C₃N₄ and NH₂-MIL-125 MOF are well matched, creating a Z-scheme heterojunction, rGO is an electron mediator crucial for suppressing charge recombination in this configuration. In 2021, Muhammad Tahir and colleagues used thermal and hydrothermal synthesis to prepare an RGO-mediated g-C₃N₄/NH₂nanocomposite (Fig. 5d).²⁰¹ This composite embedded Ti-MOF on the g-C₃N₄-RGO matrix, achieving better dispersion and enhancing charge transference. TEM images showed favorable heterojunction interactions. The neat MOF exhibited H₃ hysteresis loops, while the composite showed type H4 hysteresis due to the MOF's high surface area. Pore size distribution increased after combining MOF with g-C₃N₄, with higher pore volumes at ~50 nm. NH2-MIL-125 had intense visible-range absorption, improving solar energy use. The g-C₃N₄-RGO matrix showed improved light absorption with 20 wt% MOF loading. The g-C₃N₄-RGO-MOF (20%) composite achieved maximum efficiency with CO generation of 383.79 µmol per g, thanks to RGO's enhancement of charge carrier separation and transfer, and NH2-group functionalization in MOF increasing activity in the visible light region and also CO2 adsorption. The amino group promoted CO2 interactions, and oxygen in ligands and Ti-metal complexes increased water adsorption. At various CO2 pressures (0.2, 0.4, and 0.8 bar), the ternary composite showed increased CO (25.08 μ mol g⁻¹) and CH₄ (597.63 μ mol g⁻¹) production rates, with higher pressure enhancing CO2 concentration and availability. The indirect Z-scheme heterojunction and RGO's role as an electron mediator reduced recombination rates, boosting productivity and quantum yields.

 ${
m TiO_2}$, known for its nontoxicity, availability, and light stability, is widely used as a semiconductor. Huang *et al.* advanced photocatalytic ${
m CO_2}$ and ${
m CH_4}$ conversion using NH₂-MIL-125 with ${
m TiO_2}$ quantum dots. ¹⁴⁷ The NH₂-MIL-125@ ${
m TiO_2}$ -40 cata-

lyst showed superior CO2 conversion efficiency, achieving a CO yield of 159 μmol g⁻¹, surpassing previous reports and indicating potential for practical CO2 utilization. A heterojunction structure was found at the NH2-MIL-125 and 0D TiO2 interface in CNMT-40, while SNMT-15 exhibited excessive TiO2 crystallization, potentially affecting the MOF structure (Fig. 5e). The red-shift in CNMT-40's absorption spectrum suggests an E_{g} modification due to Ti-C-N bond formation and electronic structure changes, enhancing visible light absorption. TiO2 integration significantly improved the photocurrent density of the MOF, indicating better exciton pair separation. CNMT-40 executed superior photocurrent intensity and stable performance during tests, highlighting its superior e⁻ and h⁺ transportation and separation compared with other catalysts. The hybrid material NMT, prepared via a wet chemical method, outperformed 2D TiO2 and NM. CNMT-40 maintained a 93.7% photocatalytic CO₂ conversion efficiency after six cycles, demonstrating decent stability. Its better performance in converting CO₂ and CH₄ was due to an efficient and stable heterojunction that improved electron and hole transport rates, inhibited charge recombination, and enhanced catalyst stability. Amorphous titanium dioxide (Am-TiO2) offers benefits like milder synthesis conditions and a larger surface area.²¹⁸ Though it has a wide band gap and high recombination rates, limiting visible light harvest and thus the photocatalytic performance, Am-TiO2 can be an efficient co-catalyst. Hu et al. fabricated Am-TiO2/NH2-MIL-125 via a one-pot water bath method. 219 Am-TiO2 acted as an electron acceptor, enhancing photocarrier separation. XRD patterns showed effective integration with NM, and Raman spectroscopy confirmed the amorphous nature of Am-TiO2, showing minimal vibration modes compared with anatase TiO2. NM and NM/TiO2 composites produced significant CH₄ from CO₂ reduction, while Am-TiO₂ showed minimal activity. XRD patterns and FTIR/TGA analysis indicated no significant changes in structure or properties post-reaction, showing robustness. The 0.25 ATN composite had a stronger visible light response due to defects in amorphous TiO2, resulting in more electron-hole pairs. Generally, a smaller contact angle indicates better hydrophilicity of the catalyst.220 Contact angle tests revealed higher hydrophilicity for 0.25 ATN, enhancing photoactivity. Preirradiation bands indicated CO2 adsorption capability, while post-irradiation decreases in Ti3+-OH and Ti4+-OH peaks suggested CO2 reaction to form bicarbonate species. The mechanism involves electron photoexcitation from NM to Am-TiO2, with efficient exciton separation boosting the performance of NM/Am-TiO2 composites in CO2 reduction.

The facet-dependent photocatalytic behavior of MOFs, particularly NH_2 -MIL-125(Ti), remains largely unexplored. In one study, Cheng *et al.* synthesized NH_2 -MIL-125(Ti) with controlled {001} and {111} facet ratios *via* a one-step solvothermal method, enabling precise tuning of active facet exposure. By adjusting the solvent ratio of DMF and MeOH and varying titanium sources, they modulated the facet ratios effectively. With increased {111} facet exposure, a red shift in the absorption spectrum was observed, while NM001 displayed the narrowest

absorption. Using TEOA as a sacrificial agent, NM111 exhibited significantly higher photocatalytic performance for CO and CH₄ production, with yields of 8.25 and 1.01 μmol g⁻¹ h⁻¹—9 and 5 times higher than NM001, respectively. This enhancement was attributed to the larger surface area and higher Ti content in the {111} facets, which facilitated Ti⁴⁺/Ti³⁺ interconversion and unveiled more active sites for CO2 reduction. 221 Building on prior research, the group demonstrated that loading Ag NPs onto the {111} facets of NH2-MIL-125(Ti) via in situ photodeposition significantly enhanced photocatalytic activity and CH₄ selectivity for CO₂ reduction. The Ag-loaded catalysts exhibited more negative HOMO levels, boosting product yields. The Ag NPs improved CO2 adsorption by serving as additional adsorption sites, with a small Q_{st} value indicating weak CO2-catalyst interactions that facilitated CO2 migration to active sites. Additionally, the SPR effect of Ag NPs enhanced light absorption, charge separation, and transfer. These synergistic effects broadened light absorption, increased CO₂ adsorption, and suppressed charge recombination, offering new insights into facet engineering for optimizing CH₄ selectivity in CO₂ reduction.²²² Guo et al. developed an ultrasound-assisted method to synthesize NH2-MIL-125 nanosheets with predominantly exposed {110} facets, which contained a higher concentration of active metal clusters than {001} and {111} facets. Ultrasound waves created strong shear forces that promoted the formation of thin sheet structures. This ultrathin morphology increased active sites, enhanced charge carrier transfer, and suppressed electronhole recombination. DFT simulations revealed that in the 2D {110} facet, conduction band states were distributed across the entire outer layer, allowing more Ti atoms to participate in the catalytic reaction, thereby boosting photocatalytic performance.223

In a recent study, Cheng et al. synthesized MIL-125-NH₂(Ti) MOFs with controlled facet exposure ({001}, {110}, {111}) and mixed-facet heterojunctions ({001}/{110}, {001}/{111}, {110}/ {111}) via a solvothermal method to explore facet-dependent photocatalytic CO₂ reduction. The {110}/{111} heterojunction (T23) exhibited the highest CO and CH₄ yields, at 15.49 and 5.46 µmol g⁻¹ h⁻¹ which were 10- and 18-fold higher than the {001} facet (T1). This enhancement was due to improved charge separation and transfer facilitated by the heterojunction. DFT calculations showed that the {111} facet had the lowest Gibbs free energy barrier for *COOH formation, aiding efficient CO₂-to-CH₄ conversion.²²⁴ In a follow-up study, Cheng et al. examined high-index {112} facets of NH2-MIL-125(Ti), which significantly boosted photocatalytic activity. CO and CH₄ yields increased by 33- and 31-fold, respectively, compared with the {001}/{111} facets. The distinct surface chemistry of the {112} facets provided a larger surface area, higher CO₂ adsorption, and stronger CO2 interactions, resulting in greater adsorption enthalpy and efficient CO2 activation. Additionally, more exposed Ti-oxo clusters facilitated the Ti⁴⁺/Ti³⁺ transition, enhancing electron transfer and reducing charge recombination. This study highlights the potential of highindex facet engineering as a powerful strategy for optimizing MOF-based photocatalysts, offering a promising approach for designing next-generation MOFs with tailored facets for efficient ${\rm CO_2}$ photoreduction. 225

In conclusion, noble metals enhance charge carrier separation but are expensive. Non-noble metals are cost-effective but require precision. Hybrid and graphene-based composites improve interactions and charge transport but face challenges in uniform dispersion and stability. New strategies include using conductive polymers and Z-scheme heterojunctions. Future research should focus on optimizing material combinations, doping, facet engineering, and stability for efficient and scalable CO₂ reduction.

10.4. N₂ fixation

Ammonia (NH₃) serves as a critical precursor in fertilizer manufacturing, and industrial N_2 fixation traditionally depends on the Haber–Bosch process, which necessitates high temperature and pressure for transforming N_2 and H_2 into NH₃. Photocatalytic nitrogen fixation occurs through two main processes: nitrogen reduction reaction (NRR) and nitrogen oxidation reaction (NOR). When exposed to sunlight, the photogenerated holes in VB facilitate the oxidation of H_2O into H^+ and O_2 (eqn (15)). Additionally, NH₃ is synthesized from N_2 reduction by hot electrons under sunlight (eqn (16)), enabling NH₃ production from H_2O and N_2 under optimal conditions with solar light as the primary source of energy (eqn (17)).²²⁶

$$2H_2O + 4h^+ \rightarrow O_2 + 4H^+$$
 (15)

$$N_2 + 6H^+ + 6e^- \rightarrow 2NH_3$$
 (16)

$$2N_2 + 6H_2O \rightarrow 4NH_3 + 3O_2$$
 (17)

During photocatalytic N_2 oxidation, photogenerated holes (h^+) react with N_2 and H_2O to produce NO (eqn (18)), while photoexcited electrons reduce O_2 to H_2O (eqn (19)). The generated NO is subsequently oxidized to nitrates, resulting in the evolution of O_2 and H_2O (eqn (20)). Under ambient conditions, nitric acid is produced using sunlight as the primary energy source by synthesizing it from water, O_2 , and N_2 (eqn (21)).

$$N_2 + 2H_2O + 4h^+ \rightarrow 2NO + 4H^+$$
 (18)

$$O_2 + 4H^+ + 4e^- \rightarrow 2H_2O$$
 (19)

$$4NO + 3O_2 + 2H_2O \rightarrow 4HNO_3$$
 (20)

Overall:
$$2N_2 + 5O_2 + 2H_2O \rightarrow 4HNO_3$$
 (21)

The two pathways for photocatalytic N_2 fixation are referred to as the associative alternating route (Path 1) and the associative distal route (Path 2).²²⁷

Artificial N_2 fixation under milder conditions remains a significant challenge to date. TiO_2 containing oxygen vacancies can function as a photocatalyst for N_2 fixation, leveraging the presence of Ti(m); however, its activity is restricted to the UV region only (Fig. 5f and g). In contrast, the titanium-based NH_2 -MIL-125, featuring a titanium-oxo cluster and acting as a visible light photosensitizer, is emerging as a promising candi-

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date for achieving artificial N2 fixation under visible light conditions. Li et al. introduced a novel range of titanium-based MOFs designed for N2 fixation under visible light exposure, marking a significant advancement. 148 They synthesized three types of visible light-responsive titanium-MOFs—CH₃-MIL-125, NH₂-MIL-125, and OH-MIL-125—via either one-pot synthetic method or solvent-assisted linker exchange approaches. When compared with neat MIL-125(Ti), these MOFs modified with different ligands exhibited enhanced light-absorption capabilities, ranked as CH₃-MIL-125(Ti) < OH-MIL-125(Ti) < NH₂-MIL-125(Ti). Consequently, their photocatalytic performance also followed the same order for producing ammonia. Control experimentations employing isotope labeling confirmed N2 and water as the sources of the generated ammonia. Mechanistic insights revealed that the organic ligands, responsive to visible light, acted as efficient photosensitizers, while the reduced Ti(III) species served as the active catalytic sites. Although research about these materials is very promising, there is still a critical gap in the literature regarding the mechanistic pathways of N₂ reduction in these systems. The precise role that organic ligands play in Ti-based MOFs acting as photosensitizers is being studied, but their presence seems to be critical for achieving better electron transfer and, thus, better catalytic efficiency.

Atomic or ion doping can modify M-O clusters, thereby altering the electronic and microenvironment around these catalytic sites crucial for photocatalytic reactions. 228 Hou et al. employed a one-step solvothermal method to produce a range of Ce-doped NH2-MIL-125, aiming to investigate their efficacy in photocatalytic N₂ fixation. 149 The incorporation of Ce extends the absorption edge into the visible region and introduces additional reaction sites, thereby enhancing the charge transfer and facilitating dual electron O2 reduction reactions. Upon light excitation, the Ce⁴⁺/Ce³⁺ redox pair adjusts e-states adjacent to the Fermi level, engaging additional active e-s in N₂ fixation.²²⁹ Ce doping also increases the Ti³⁺ content, which serves as an active site and facilitates N2 chemisorption by donating electrons. Additionally, Ce boosts LMCT efficiency, improving the separation of photogenerated carriers. Under full-spectrum radiation, NH2-MIL-125 achieves an NH₃ generation rate of 10.6 μmol g⁻¹ h⁻¹. Incorporating 2% Ce into the MOF matrix significantly enhances its N2 fixation performance, achieving an NH3 synthesis rate of approximately 39.4 μ mol g⁻¹ h⁻¹, which is 3.7-fold better than the undoped MOF. This improvement underscores the synergistic effect between Ti⁴⁺/Ti³⁺ and Ce⁴⁺/Ce³⁺ ions, which boosts the proficiency of the photocatalytic nitrogen fixation reaction. While doping improves the catalytic efficiency, the relative success of doping depends on finely tuning the dopant concentration to avoid loss of structural integrity or unwanted secondary reactions in real-world situations. The main challenge in nitrogen fixation is achieving high catalytic efficiency under mild conditions by optimizing heterojunctions for improved exciton separation and light absorption. Ensuring stability and reusability during operation is vital, particularly when adding dopants to enhance activity and maintain structural integrity.

Mechanistic studies of electronic interactions are crucial for designing more potent catalysts.

10.5. H₂ evolution

Photocatalytic H_2 production via water splitting represents a potential solution to the global energy crisis. 52,230,231 The process of photocatalytic H_2 evolution involves three main stages: (i) photons with energy exceeding the semiconductor's E_g excite e^-s from the VB to the CB, simultaneously generating h^+s in the VB; (ii) the exciton pairs generated get separated and migrate to the exterior of the material; (iii) surface adsorbed H^+ is reduced by the e^-s in the CB to form H_2 , while holes in the VB oxidize H_2 O to produce O_2 . For efficient H_2 production, the following redox reactions must occur: 232

Oxidation:
$$2H_2O \rightarrow O_2 + 4H^+ + 4e^ E_0 = 1.23 \text{ eV}$$
 (22)

Reduction:
$$2H^+ + 2e^- \rightarrow H_2$$
 $E_0 = 0.00 \text{ eV}$ (23)

Overall:
$$2H_2O \rightarrow O_2 + 2H_2$$
 $E_0 = 1.23 \text{ eV}$ (24)

For HER to occur, the CB potential must be lower than 0 V under the NHE. Additionally, the VB potential must be higher than 1.23 V for the water oxidation reaction to proceed.

Since Matsuoka et al.'s pioneering work in 2012 using NH2-MIL-125 as a photocatalyst for H2 generation, numerous titanium-based MOFs have been developed for this reaction. Pt metal particles are commonly used as efficient cocatalysts. Pt NPs photodeposited on the MOF enhanced its activity, yielding a visible region active photocatalyst PtNPs@NH2-MIL-125(Ti) with a hydrogen generation activity of 367 μmol g⁻¹ h⁻¹, using e donor triethanolamine (Fig. 6a). 6 Mechanistic studies have revealed that upon excitation, photogenerated electrons on NH₂-BDC transfer initially to the metal-oxo cluster and subsequently to platinum NPs towards H2 production. UV light illumination has been shown to significantly enhance activity, reaching 1170 μmol g⁻¹ h⁻¹, primarily because of the direct excitement of titanium-oxo clusters. Additionally, Matsuoka et al. explained the improved photoactivity of PtNPs@NH2-MIL-125 by optimizing the platinum deposition. They determined that a platinum loading of 1.5 wt% achieved the maximum activity of 517 µmol g⁻¹ h⁻¹. Besides the outstanding performance of the PtNPs@NH2-MIL-125(Ti) system, a major drawback seen in this system is the structural instability and Pt NP leaching during long-term reactions.

In response to this challenge, C. Stylianou and team explored the photocatalytic $\rm H_2$ evolution of $\rm Pt/NH_2\text{-}MIL\text{-}125$ composite materials functionalized with three distinct ligands: $(\rm NH_2)_2\text{-}BDC$, OH-BDC, and $(\rm OH)_2\text{-}BDC$. Without Pt NPs, hydrogen generation rates ranged between 39 and 80 μ mol h⁻¹ g⁻¹. Adding Pt NPs improved performance, with optimal loading of about 2.5 wt%. Pt/NH₂-MIL-125/(OH)₂ exhibited the highest activity at 707 μ mol h⁻¹ g⁻¹, followed by NH₂-MIL-125 at 619 μ mol h⁻¹ g⁻¹, while NH₂-MIL-125/OH and NH₂-MIL-125/(NH₂)₂ had lower rates (377 and 253 μ mol h⁻¹ g⁻¹, respectively). Despite NH₂-MIL-125/(NH₂)₂'s reduced E_g suggesting enhanced visible-light absorption, its photo-

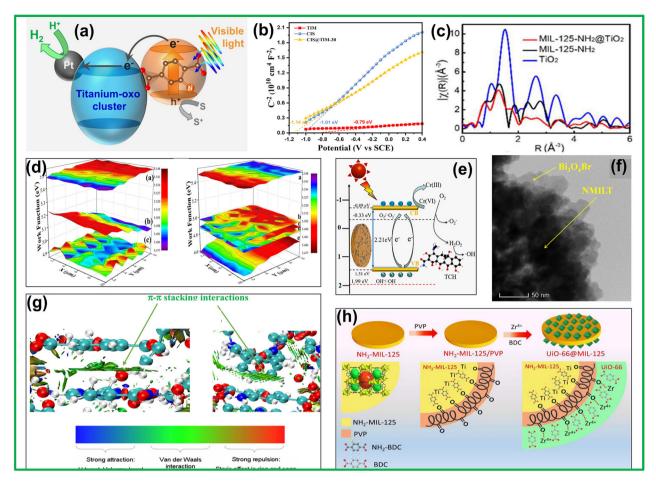


Fig. 6 (a) A schematic illustration of photocatalytic hydrogen production reaction over Pt-supported Ti-MOF-NH₂ based on the LCCT mechanism. Reprinted with permission from ref. 56. Copyright 2012 American Chemical Society. (b) Mott-Schottky curve of pure CIS, TIM, and CIS@TIM-30. Reprinted with permission from ref. 150. Copyright 2022 Elsevier Ltd. (c) EXAFS spectra at Ti K-edge. Reprinted with permission from ref. 158. Copyright 2018 American Chemical Society. (d) The work functions of ZIS, TiM, and CdS (left), respectively, and ZIS, TiM/CdS, ZIS/CdS, TiM@ZIS, and TiM@ZIS/CdS (right), respectively. Reprinted with permission from ref. 154. Copyright 2019 Elsevier B.V. (e) Diagram showing an example of the photocatalytic mechanism of NMTi-OVs-300. Reprinted with permission from ref. 159. Copyright 2023 Elsevier B.V. (f) TEM image for BOB/NMILT composite. Reprinted with permission from ref. 160. Copyright 2022 Elsevier B.V. (g) 3D Plot of noncovalent interactions ($\pi-\pi$ stacking) using an NCI analysis technique for MIL-125 NH₂ interaction with caffeine (left) and carbamazepine (right). Reprinted with permission from ref. 162. Copyright 2022 Elsevier B.V. (h) Schematic representation for the systematic growth of MOF-on-MOF heterostructure. Reprinted with permission from ref. 95. Copyright 2022 Elsevier B.V.

catalytic performance was inferior. Computed wavefunctions revealed that NH₂-MIL-125/(OH)₂ had greater charge carrier mobility due to the spatial distribution of electron density within Ti-octahedra and NH2-BDC ligands, leading to longer lifetimes of photogenerated charge carriers. This enhanced charge separation and migration improved photocatalytic efficiency. NH₂-MIL-125/(OH)₂ displayed the highest conductivity $(5.79 \times 10^{-6} \text{ S m}^{-1})$, attributed to increased electron density, enhancing redox potential and electron spatial distribution. The effects of ring activation of -OH groups were found to be greater than those of amines, enhancing p- π conjugation and electron conductivity, thereby improving photocatalytic efficiency. Such studies relate ligand activation to charge mobility but may require further research to optimize the interaction between the electronic properties of the ligand and the metal-oxo clusters to achieve greater efficiency. Charge

mobility was increased with NH₂-MIL-125/(OH)₂, but performance remained constrained by recombination in the bulk material, thus demonstrating the need for synergistic ligand engineering with prevention of bulk-recombination strategies.

Zhang et al. fabricated CdS/NH2-MIL-125 Z-scheme heterojunction via a solvothermal process. 151 FESEM images showed CdS nanoparticles adhering to NH2-MIL-125 nanoplates, with fewer nanoparticles as Ti-MOF concentration increased, promoting better contact interfaces for charge carrier migration. TEM confirmed strong interfacial contact between CdS NPs and MOF nanoplates. The absorption edge at 517 nm corresponds to CdS's band gap, while bands at 280 nm and 370 nm are due to metal oxo clusters and amine ligands. CM-X samples absorbed less visible light with more Ti-MOF loading. Photocurrent intensity improved with NH2-MIL-125(Ti), except for CM-50, likely due to the "shielding effect". CM-10 showed

the highest charge separation and migration efficiency, with the strongest photocurrent signal and an $\rm H_2$ production rate of 6.62 mmol $\rm h^{-1}g^{-1}$, about 3.5-fold more than CdS alone. ESR experiments under simulated sunlight showed characteristic $\rm O_2^{\, '-}$ radical signals, indicating a Z-scheme heterojunction with efficient electron transfer. Pure $\rm NH_2\text{-}MIL\text{-}125(Ti)$ and CdS showed weaker signals, and 'OH radical experiments supported the Z-scheme hypothesis, with stronger signals for Ti-MOF and CM-10 compared with pristine Ti-MOF. The absence of 'OH radicals in Type-II heterojunctions supported the Z-scheme heterojunction hypothesis.

Ternary metal sulfides like CdIn₂S₄ (CIS) are notable for narrow band gaps and strong visible absorption. 233,234 NH₂-MIL-125 is effective for H₂ production due to its porous nature and H₂ adsorption, ¹⁵⁴ but its stability and efficiency are compromised by moisture and high recombination rates. 235,236 To address this issue, Zhou et al. created a CdIn₂S₄@NH₂-MIL-125 core-shell structure via hydrothermal methods for H_2 generation. This heterojunction enhanced carrier separation and migration, improving photocatalytic performance by matching band gaps and extending carrier life. Increased CdIn₂S₄ content raised peak intensity, confirming incorporation, while NH2-MIL-125's absence in diffraction peaks indicated its encapsulation. Evaluated under visible light with Na₂S/Na₂SO₃ as a sacrificial agent, CIS showed a steady hydrogen production increase, reaching 2650 µmol g⁻¹ in 3 hours due to efficient carrier separation at the CIS-TIM heterojunction. Higher CIS loading didn't always improve performance, likely due to excess CIS masking active sites. CIS@TIM-30 exhibited the highest performance, producing 7650 µmol g⁻¹ hydrogen in 3 hours, over three times that of pure CIS. Mott-Schottky (M-S) curves (Fig. 6b) confirmed the n-type behavior of both CIS and TIM, with CIS@TIM-30 showing enhanced charge separation efficiency. The lower arc radius in the EIS study and lower PL intensity of CIS@TIM-30 indicated improved charge separation and accelerated electron transfer at the CIS-TIM heterojunction. Efficient electron transfer from TIM to CIS enables hydrogen evolution through H⁺ reduction in water. With optimized carrier dynamics and a robust heterojunction interface, CIS@TIM-30 demonstrates superior photocatalytic performance. Tianding Hu and colleagues, using a hydrothermal method, developed a novel ternary photocatalyst, NH2-MIL-125(Ti)/CdIn2S4/Pd, where Pd was incorporated into NH₂-MIL-125(Ti). The optimized Pd/ NH₂-MIL-125(Ti)/CdIn₂S₄ material displayed a remarkable photocatalytic H₂ evolution activity of 4.60 mmol g⁻¹h⁻¹. This rate was 31 times greater compared with neat CIS and 16 times more than the parent MOF. The boosted activity resulted from efficient energy band alignment between CIS and Ti-MOF, facilitating electron transfer under visible light. Pd nanoparticles create active sites for H₂ production, reducing electron-hole recombination. This enhances light absorption, improves mass transfer, and offers more active sites, significantly improving catalytic performance.

Liu and team synthesized a series of $ZnIn_2S_4@NH_2-MIL-125$ heterostructures with varying MOF contents using a

straightforward solvothermal approach. 153 The nanosheets of ZnIn₂S₄ were well dispersed on the MOF matrix, and the composite photocatalysts demonstrated superior photocatalytic action compared with their individual components under visible-light conditions for H2 generation. The incorporation of NH₂-MIL-125 MOF significantly enhanced the SSA and pore volume of the composite, particularly at higher MOF content, resulting in a higher number of active sites for increased adsorption of the substrate. The ideal MOF content was approximately 40 wt%, resulting in a photoinduced hydrogen production rate of 2204.2 μ mol h⁻¹g⁻¹ and an AQE of 4.3% at 420 nm. This performance was 6.5 times better than that of neat ZIS. The superior photocatalytic performance can be credited to the favorable band alignment and the close interfacial contact interface between ZIS and the MOF, which enabled photogenerated charge carriers' efficient transfer and separation. Considering these factors, ZIS and CdS were used to create an ordered tandem heterojunction, with CdS as a co-catalyst on the titanium MOF core and ZIS shell, enhancing charge carrier separation. 154 Fig. 6d illustrates the work function maps identifying samples labeled as ZIS, TiM, and CdS for (a), (b), and (c), respectively. The work functions of (a), (b), and (c) measured approximately 5.53, 5.15, and 4.93 eV, respectively. This indicates the hierarchy of electron transfer abilities in TiM@ZIS/CdS heterostructure as follows: ZIS > TiM > CdS. Therefore, electrons were observed to transfer from TiM to ZIS and from CdS to ZIS, validating the electron enrichment strategy. This configuration resulted in a significantly enhanced photocatalytic hydrogen production rate, reaching 2.367 mmol g⁻¹h⁻¹. This innovative core@shell hierarchical tandem heterojunction provides valuable insights for developing high-functioning heterojunctions with enhanced multi-channel charge transference capabilities.

In Z-scheme photocatalysts, close semiconductor contact is vital for effective charge separation, with mediators like MXenes improving this by forming Schottky barriers and reducing noble metal shielding. Schottky barriers and reducing noble metal shielding. Liu *et al.* synthesized a Z-scheme photocatalyst (NH₂-MIL-125, Ti₃C₂ MXene QDs, and ZnIn₂S₄) for visible-light-induced H₂ production, achieving a high rate of 2931.9 μ mol g $^{-1}$ h $^{-1}.^{155}$ Although not all Ti₃C₂ MXene QDs bridge both semiconductors, some facilitate charge transfer in a Z-scheme pathway. Superoxide trapping and redox potential tests suggest a Z-scheme mechanism, where electrons transfer from Ti-MOF's CB to ZnIn₂S₄'s VB *via* Ti₃C₂ MXene QDs.

1T-MoS₂, an economically viable and abundantly available metal sulfide, is recognized as a promising co-catalyst.²⁴¹ Nguyen *et al.* established a photocatalyst nanocomposite by integrating 1T-MoS₂ into NH₂-MIL-125, showing superior rates of photocatalytic hydrogen production compared with Pt/NH₂-MIL-125, Co-oxime@NH₂-MIL-125, and Ni₂P/NH₂-MIL-125. The higher output of 1T-MoS₂ was ascribed to its plentiful reaction sites present on both the edges and the basal planes, which contrasts with 2H-MoS₂, which predominantly features active sites only on its edges.^{156,242}

Ti-MOF composites assisted by non-noble metals have shown effectiveness in photocatalytic water splitting and hydrogen evolution. Nasalevich et al. pioneered a cobaloximederived Ti-MOF photocatalyst that established a more noteworthy 20-fold rise in H2 evolution rise than NH2-MIL-125. 123 The catalyst exhibited excellent stability, maintaining a constant turnover frequency even after 65 hours of light irradiation, and retained its efficiency through multiple cycles of experiment. The photocatalytic activity of Co@NH2-MIL-125 was found to be less than the LUMO of the neat MOF. In this system, photoinduced h⁺s located at the organic ligand interact with a sacrificial e donor, while electrons are promptly transferred to cobalt species, reducing Co(III) to Co(II) and thereby enhancing overall photocatalytic activity due to the existence of high-spin Co(II) species. Recently, Cu²⁺/Cu⁺ was introduced into the linkers to form mixed-valence redox sites, leading to a substantial increase in charge density by 7000 and lifetime by 27 times, respectively. This modification led to a remarkable improvement in the hydrogen production rate of the Cu-modified NH₂-MIL-125(Ti) to 490 mmol g⁻¹ h⁻¹, which represents an increase of 27 times compared with parent MOF and 10 times than Pt-NH₂-MIL-125(Ti).²⁴³ In 2016, Meyer and colleagues integrated the Ni(II) catalyst [Ni(dmobpy) (2-mpy)₂] into NH₂-MIL-125, leading to the development of Ni@NH2-MIL-125(Ti), marking the first among MIL-based photocatalyst to incorporate Nickel(II) particles.²⁴⁴ Under light irradiation, the composite system demonstrated hydrogen yields that were 1800-fold higher than bare MOF and ten times higher than the Ni catalyst molecular catalyst. Moreover, the prepared nanocomposite exhibited a 10-fold increase in catalytic activity and sustained performance for 3000 minutes. This increase was attributed to mixed-valence redox sites, which enhanced the charge density and carrier lifetime. The main challenge related to non-noble metal co-catalysts is stability, with efficient charge transfer persisting over time. Their electronic properties must be tuned towards the photocatalytic process, with long-term stability, for them to compete with platinum.

A variety of co-catalysts based on transition metals (Ni₂P, NiO, Co₃O₄, CoP, CuO, Fe₂O₃) were united with the Ti-MOF, leading to enhanced visible light-induced photocatalytic H₂ production. Remarkably, NiO/NH2-MIL-125(Ti) and Ni2P/NH2-MIL-125 demonstrated notable H2 generation rates of 1084 and 1230 µmol g⁻¹ h⁻¹, respectively. 157 PL study indicated that all the co-catalysts can suppress the emission of neat MOF, with the efficiency of charge separation ranked as Ni₂P > CoP $\approx \text{Co}_3\text{O}_4 > \text{Fe}_2\text{O}_3 > \text{CuO} > \text{NiO}$. This order was in agreement with the photocatalytic H2 evolution results, except for NiO/ NH₂-MIL-125. Cyclic voltammetry measurements addressed this discrepancy, showing that NiO had a current density around three times higher than Ni_2P at -0.9 V versus Ag/AgCl, indicating higher intrinsic activity of NiO that partly compensates for its lower electron capture ability.

Zhang et al. developed a post-solvothermal method for removing organic ligands from MOFs, providing a simple, fast process at ~350 °C compared with conventional pyrolysis. 158

This method created NH2-MIL-125 NPs coated with TiO2 nanosheets, forming a core-shell type NH2-MIL-125@TiO2 structure. This structure combined active TiO2 nanosheets, MOF photosensitizers, linker deficiency, oxygen vacancies, and a mesoporous architecture, enhancing H2 production under visible light exposure. The process reduced the BET surface area from 1169 m² g⁻¹ to 780 m² g⁻¹ due to linker etching. EXAFS spectra showed deviations from individual NH2-MIL-125 and TiO2, indicating changes in metal site coordination, with peak shifts and reductions suggesting linker defects and oxygen vacancies advantageous for catalytic processes, enhancing active sites and facilitating the efficiency of mass/charge transfer (Fig. 6c). Varying temperatures and solvents affected the core-shell structure formation: propanol yielded core-shell particles, while methanol resulted in multiporous structures. The NH₂-MIL-125@TiO₂ composite boosted photocatalytic hydrogen production by enhancing titaniumoxo cluster exposure, light harvesting, and active sites, outperforming the parent MOF. It also showed durability across three runs, maintaining its crystallinity and morphology due to the protective TiO2 layer.

The introduction of suitable co-catalysts can further enhance the photocatalytic output of MOF/g-C₃N₄ hybrids. In 2015, Su et al.⁵⁷ conveyed the successful preparation of NH₂-MIL-125/Pt/g-C₃N₄ hybrids, demonstrating a notable threefold enhancement in photocatalytic hydrogen production compared with Pt/g-C₃N₄ alone. This approach highlights the synergistic benefits of integrating an appropriate co-catalyst with robust light-absorption capacity, leveraging both interface effects and localized SPR to augment the productivity of light energy conversion for H₂ generation. ^{245,246} These co-catalysts incorporate metal nanoparticles that efficiently absorb light, thereby enhancing visible-light absorption, generating more exciton pairs, and facilitating energy transfer to the semiconductor's CB. Additionally, synergistic effects in various photocatalytic applications have been demonstrated by combining two different metal nanoparticles (NPs).²⁴⁷ Xu et al. expanded on this concept by anchoring CN nanosheets with NH2-MIL-125 and subsequently loading Ni/Pd co-catalysts onto the NH2-MIL-125/CN surface to create an NH2-MIL-125(Ti)/CN/NiPd nanostructure. 248 Ni/Pd co-catalysts, as photosensitizers, enhanced light absorption and photocatalytic productivity of CN and MOF. The close interface between MOF and CN enabled direct charge transference and efficient charge carrier separation. Consequently, the resulting composites demonstrated a significantly elevated hydrogen production rate of 8.7 mmol h⁻¹g⁻¹, 322 times greater than NH₂-MIL-125(Ti)/CN alone. Upon discussion, it is realized that while the noble metal loading onto the MOF can enhance hydrogen production, issues like structural collapse and metal leaching persist. Key issues involve improving catalyst stability, charge carrier separation, and reduction of dependence on precious metal co-catalysts. A comprehensive approach combining material design, catalytic optimization, and stability will be necessary to address these challenges.

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10.6. Drug degradation

The photocatalytic degradation of pharmaceuticals involves three primary steps: photon absorption, excitation, and reaction.⁵³ The photocatalyst absorbs photons, generating exciton pairs that move to the surface to react with adsorbed substances. Photogenerated holes can directly break down drug molecules (eqn (26)). Two degradation pathways exist: a reductive pathway, where electrons react with O2 to form superoxide radicals (O2 *-) if the CB potential is more negative compared with O_2/O_2 redox potential (-0.13 eV vs. RHE) (eqn (27)), and an oxidative pathway, where holes generate hydroxyl radicals ('OH) from H₂O/OH⁻, depending on pH (eqn (28)). Recombination of hydrogen ions with electrons generates heat, reducing efficiency (eqn (29)). The photocatalyst's redox potential must exceed 'OH/OH" (+1.99 eV vs. RHE). 249,250 Reactive radicals ('OH and O2'-) efficiently mineralize drugs and intermediates into water and CO2 under prolonged UV exposure (eqn (30)). 251,252 Both pathways synergistically prevent electron accumulation in the CB and reduce electronhole recombination compared with direct interaction pathways.

Photocatalyst
$$+ hv \rightarrow \text{photocatalyst} + h^+ + e^-$$
 (25)

$$h^+ + drug \rightarrow H_2O + CO_2 + degradation products$$
 (26)

$$O_2 + e^- \rightarrow O_2^{\bullet -} \tag{27}$$

$$H_2O/OH^- + h^+ \rightarrow {}^{\bullet}OH + H^+$$
 (28)

$$H^+ + e^- \rightarrow energy$$
 (29)

Drugs + 'OH or
$$O_2$$
' $^- \rightarrow CO_2 + H_2O + degradation product$
(30

Introducing defect sites enhances charge separation, stability, and performance by preserving the crystal structure and increasing active sites. Vacant O-sites trap electrons, speeding up exciton pair generation and separation, boosting charge carrier generation, and improving reactant adsorption and activation. This alters photoelectric properties, enhancing photocatalytic activity. Li et al. developed NH2-MIL-125 MOF enriched with vacant O-sites to eliminate tetracycline hydrochloride and hexavalent chromium with simulated sunlight. 159 They introduced oxygen vacancies through heat treatment and examined how varying the treatment duration affected the SSA, band configuration, vacant sites' concentration, mesoporous structure, and the efficiency of exciton pair migration and parting in the oxygen-enriched NMTi-OVs-X photocatalyst. Increased heat-treatment time led to the gradual decomposition of the MIL linker and penetration from the exterior to the interior, enhancing oxygen vacancy amount. The presence of optimally vacant oxygen in NMTi-OVs-X improved the photogenerated charge carrier separation, enhancing the catalytic performance. These vacancies, activated by O2 to form reactive species (such as 'O2, 'OH, and H2O2), improved pollutant degradation efficiency. The oxygen-enriched Ti-MOF achieved 100% photocatalytic elimination efficiency of Cr(vi) and TCH

degradation in 2 hours under simulated sunlight with minimal heat treatment at 300 °C for 300 minutes (Fig. 6e). This demonstrates the potential of single-MOF strategies for environmental remediation applications.

To improve the activity of pristine MOFs and address their limitations, hybridizing semiconducting MOFs with another semiconductor to create heterostructures is considered an effective strategy. This approach is thought to combine the advantages of both materials - e.g., enhanced charge carrier separation, improved adsorption properties, and synergistic catalytic effects – for higher photocatalytic efficiency. Hu et al. fabricated a microspherical Bi₃O₄Br/NH₂-MIL-125(Ti) (BOB/ NMILT) composite photocatalyst, where Bi₃O₄Br nanosheets were firmly wrapped over the NMILT surface using a hydrothermal method (Fig. 6f). 160 The optimal BOB/NMILT-5 exhibited significant photocatalytic activity, achieving 88.5% tetracycline (TC) degradation in 90 minutes under simulated solar light, 6.5 and 3-fold superior to parent MOF and Bi₃O₄Br, respectively. This enhancement was primarily ascribed to the Z-scheme heterojunction construction, which effectively minimized exciton charge recombination, and the major role played by radicals such as h⁺ and 'O₂ in the TC degradation pathway. The building of hybrid materials based on MOF and binary semiconductor systems with layered nanostructures has significantly improved. Muelas-Ramos studied the photocatalytic removal of diclofenac using hybrid materials fabricated from g-C₂N₄ and NH₂-MIL-125, varying the loading amounts (MOF: C_3N_4 ratios of 25:75, 50:50, and 75:25). 161 Analysis of morphological features revealed plate-like particles and fiber-like forms from the MOF and g- C_3N_4 . The E_g of the hybrid materials was similar to the parent MOF, indicating that the optoelectrical properties depend on both components. The synergistic effect of the MOF/g-C₃N₄ (50:50) hybrid resulted in excellent photocatalytic activity, achieving complete diclofenac degradation in 2 hours under LED irradiation. Optimizing the interface between the two materials is one of the major hurdles in this approach. Even though the concept of heterojunction is theoretically attractive, the interface quality is of great importance for efficient charge transfer and overall photocatalyst efficiency.

MOFs and their heterostructure composites with perovskites have been utilized in photocatalysis, with NH2-MIL-125 MOFs serving as a stable, highly photoactive matrix for incorporating LaFeO3. Younes et al. studied the composite's efficacy for photocatalyzing carbamazepine and caffeine degradation. 162 The adsorption mechanism at the interface among pollutants and NH₂-MIL-125, LaFeO₃ involves metal- π , π - π stacking, and H-bonding relationships (Fig. 6g). The optimal LaFeO₃/NH₂-MIL-125 composite exhibited excellent degradation efficiencies (carbamazepine 74%, caffeine 87%) due to efficient heterojunctions facilitating charge separation. MOFon-MOF heterostructures, formed by growing one MOF on another, showed outstanding photocatalytic performance in aqueous media thanks to their structural stability, enhanced interfacial contact, large specific surface area, and increased porosity, advancing pollutant degradation. Sepehrmansourie

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S-scheme charge transfer and improving redox performance. The hollow sandwich-like structures with exposed incremental active sites contributed significantly to the improved photocatalytic activity for TCH degradation.

et al. fabricated NH2-MIL-125@UiO-66, a Z-scheme heterojunction, by epitaxially growing UiO-66 on the surface of MOF using a solvothermal technique, creating MOF-on-MOF heterostructure (Fig. 6h).95 The hybrid was then decorated with g-C₃N₄ nanosheets to design a dual Z-scheme heterojunction, UiO-66/NH₂-MIL-125/g-C₃N₄ photocatalyst for visible lightinduced OFL degradation. Constructing MOF-on-MOFs with different physicochemical structures is challenging due to the PVP and Zr⁴⁺ ions' polar interactions in UiO-66, enhancing UiO-66's affinity for NH2-MIL-125. Kitagawa et al. found PVP's zeta potential to be -0.27 mV and UiO-66's to be +44.0 mV, indicating greater affinity between UiO-66 and PVP/NH2-MIL-125. Consequently, UiO-66 nuclei accumulated over NH2-MIL-125 crystals, forming micelles, followed by internal growth via solvothermal conditions. Epitaxial growth of UiO-66 over NH₂-MIL-125's edge was attributed to the {110} NH₂-MIL-125/{001} UiO-66 lattice similarity, as confirmed by TEM and EDX spectrum data. Decorating g-C₃N₄ nanosheets on the exterior surface of the MOF-on-MOF matrix preserved both MOFs' structural and morphological aspects but reduced NH2-MIL-125 crystal size (200-250 nm) while increasing UiO-66 crystal size (230-300 nm). The double Z-scheme heterojunction significantly improved photocatalytic performance for OFL degradation under visible light. Photocatalytic mechanism studies, scavenging tests, and EPR techniques established the formation of 'OH and O2' radicals, crucial for OFL photodegradation. This study provides an updated, straightforward approach for MOF-on-MOF architecture and dual Z-scheme heterojunction materials, advancing environment and energyrelated applications. MOF-on-MOF heterojunctions improve interfacial charge transfer efficiency, enhancing photocatalytic performance while preserving the porous structure for superior adsorption. Constructing such a heterojunction allows the two MOFs to maintain their porous nature, thus boosting adsorption capabilities. For example, Gao et al. developed a Z-scheme heterojunction of NH2-MIL-125@MIL-88B by solvothermal growing of MIL-88B on a MOF matrix. 164 This heterojunction leveraged the advantages of both MOFs, providing atomic-level charge transport channels, promoting directional charge transfer, and significantly improving tetracycline degradation using visible light, achieving a 96.26% degradation rate within 120 minutes. Different from conventional MOF and non-MOF hybrids, the groundbreaking MOF-on-MOF strategy in S-scheme heterojunctions combines the benefits of double semiconducting units of MOF into a single unit through a close interface connected by well-defined chemical bonds, enhancing photocatalytic properties. Yuan et al. created a NH₂-MIL-125@CoFe Prussian blue analog (PBA) heterojunction that operates on an S-scheme mechanism. 163 This MOF-on-MOF composite structure has a sandwich-like structure, having hollow CoFe PBA nano cages precisely brought together on the upper and lower surfaces of Ti-MOF. Theoretical and experimental discoveries confirmed the establishment of an internal electric field through interfacial Ti-O-Co bonds within the heterojunction. This internal field acted as a driving force and an atomic transport pathway, facilitating

Beyond heterostructure design, incorporating noble metals (e.g., Au, Ag, Pt) into MOFs is highly effective in enhancing semiconductor photocatalyst performance. Noble metal/MOF hybrids benefit from (i) the surface plasmonic resonance (SPR) effect of noble metal nanoparticles, broadening optical responses in the visible range, and (ii) the establishment of Schottky junctions with other semiconductors, controlling photoinduced electron flow and facilitating exciton pair separation. In this context, Zheng et al. fabricated Ag/NH2-MIL-125(Ti)/CdS photocatalysts, demonstrating high catalytic activity for photodegrading ketoprofen (KP) under simulated sunlight (94.2% after 180 min). 165 The hybrid material exhibited enhanced visible light harvesting and robustness compared with the neat MOF or CdS alone. This improvement in photocatalytic performance was largely attributed to the Z-scheme heterojunction activity and efficient hot e conduction facilitated by Ag nanoparticles. The material achieved a remarkable 94.2% degradation of potassium persulfate (KP) after 180 minutes under simulated sunlight conditions. Following this, Muelas-Ramos reported the solvothermal preparation method for depositing NH₂-MIL-125 with Pd, Pt, and Ag nanoparticles, aiming to maintain visible light photocatalytic behavior. 100 The incorporation of noble metals did not induce structural changes; however, it did alter the surface properties, which varied depending on the specific noble metal used. XPS characterization indicated the reduced state of the metal nanoparticles (1.8 to 3.8 nm) in the hybrid $(M/NH_2-MIL-125, where M = Pd, Ag, Pt)$, enhancing visible range activity as a result of the SPR effect of the noble metal NPs. The presence of these nanoparticles in the MOF minimized charge recombination, verified by photoluminescence techniques. Pt/NH2-MIL-125 exhibited the maximum photocatalytic activity, achieving maximum acetaminophen (ACE) conversion in less than 3 hours under simulated solar irradiation. Yet, effective synthesis of the MOFbased composite has to be done to avoid aggregation of noble metal nanoparticles, which would decrease its efficiency. Another important direction that requires more attention is selective photocatalytic degradation. Generally, photocatalysts degrade a variety of contaminants, but being selective toward a specific drug or pharmaceutical would be helpful in cases where certain contaminants need to be removed from the water.

10.7. Dye disposal

The disposal of dyes from industrial wastewater into natural water bodies poses serious environmental and health risks. Current removal methods often fail to fully degrade dyes and generate excess waste. Among the conventional water treatment methods, photocatalysis offers a promising solution for efficiently degrading toxic organic pollutants in wastewater while minimizing waste. The detailed photocatalytic

mechanism can be elucidated through the following reactions 253-255

$$H_2O + h\nu \rightarrow HO^{\bullet} + H^{\bullet}$$
 (31)

$$HO' + O_2 \rightarrow HO_2'$$
 (32)

$$O_2 + e^- \rightarrow O_2^{\bullet -} \tag{33}$$

$$O_2^{\bullet -} + H^+ \rightarrow HO_2^{\bullet}$$
 (34)

$$HO_2$$
 \cdot $+ H^+ \rightarrow H_2O_2$ (35)

$$H_2O_2 \xrightarrow{h\nu} 2HO$$
 (36)

Ultimately, the dye undergoes a reaction with a hydroxyl radical (HO'), leading to the formation of intermediate and final products:

$$HO^{\:\raisebox{3.5pt}{\text{\circle*{1.5}}}} + dye \to Intermediates \to CO_2 + H_2O$$

MOF/CdS composites are effective in photocatalysis but suffer from instability and rapid charge recombination, while noble metals enhance charge transfer through SPR effects in heterogeneous architectures.²⁵⁶ Wang et al. innovatively produced a ternary nanocomposite photocatalyst with extended visible range responsivity and heightened absorption efficiency by sequentially loading Au and CdS onto NH2-MIL-125 using a dual-step approach. 166 The CdS-core and Aushell structure not only stabilized CdS but also utilized the SPR effect of Au to facilitate efficient carrier separation, leveraging the compatible energy band alignments. Experimental zeta potential analyses verified the presence of electrostatic interactions, confirming the successful fabrication process 7a). Compared with pristine CdS, the 5Au@40NMT/CdS system exhibited superior degradation efficiency of 93.3%- and nine-times faster degradation kinetics than CdS within just 30 minutes. This enhancement stemmed from the expansive SSA of the MOF, which facilitated the dispersion of CdS and Au nanoparticles, thus encouraging efficient carrier separation. The visible light harvesting ability of CdS and enhanced transference of carriers can be achieved by the incorporation of Au NPs. ESR analysis identified superoxide radicals as the primary entity accountable for the elimination of MB dye.

Graphene is an exceptional 2D material showcasing high thermal stability and charge carrier mobility with remarkable SSA, making it an ideal support for photocatalysts. Nivetha et al. presented an NH2-MIL-125(Ti)/CdS/Graphene ternary composite that exhibited effective catalytic activity for exciton pair separation, low overpotential, and better activity in comparison with parent MOF. 167 The light-responsive behavior was studied, where NH2-MIL-125(Ti) and CdS/Graphene both showed n-type semiconductor behavior. NH2-MIL-125(Ti)/CdS/ Graphene (5 wt%) demonstrated superior charge separation efficiency. The TGA analysis results showcased significant weight loss due to solvent removal and MOF decomposition, followed by the production of amorphous TiO2 residue (Fig. 7b). The NH₂-the ternary composite (5 wt%) exhibited

enhanced thermal stability due to the presence of graphene. NH₂-MIL-125(Ti) and NH₂-MIL-125(Ti)/CdS/Graphene (20 wt%) composites had surface areas of 1156 m² g⁻¹ and 544.78 m² g⁻¹, respectively. The lower surface area of NH₂-MIL-125(Ti)/ CdS/Graphene (20 wt%) than the parent MOF is because of the internal cavities occupied by the CdS/Graphene surface.

MoS2, with its layered structure, promotes carrier separation, enhancing the adsorption of e and photoactivity, especially when narrow bandgap MoS2 (1.9 eV) forms heterojunctions with other semiconductor catalysts. 257-259 Zhang et al. reported a ternary composite of MoS₂ nanosheets, ultrasmall titanium-based MOF nanodots (<10 nm), and CF matrix as a competent and recyclable photocatalyst. 168 ROS have restricted diffusion distances (20 nm), 260 making ultra-small nanodots more effective in generating ROS under visible light for pollutant photodegradation. The ternary composite exhibited both type I and IV isotherms related to microporous and mesoporous nature, enhancing adsorption capacity and catalytic efficiency. Under the visible light illumination, the ternary composite cloth efficiently eliminated 81.1% of LVFX drug and 94.3% of MB dye, attributed to the ultra-small NH₂-MIL-125 nanodots being decorated onto the CF/MoS2 matrix (Fig. 7c). The CF/MoS₂/NH₂-MIL-125(Ti) nanocomposite showed a higher surface area of 290.1 m² g⁻¹, supplying active sites for pollutant binding. Large MoS2 nanosheets enhanced nanodot growth and created mesopores for better pollutant removal through synergistic effects and effective ROS formation. The nanocomposite cloth $(4 \times 4 \text{ cm}^2)$ remained stable and effective after four recycling cycles (Fig. 7d). This study addressed the potential of nanostructuring with layered materials to significantly improve photocatalytic performance, an approach that has great prospects in recyclable photocatalysts that can be readily used without losing efficiency.

Various bismuth-based metal oxide photocatalysts have been developed to enhance photocatalytic efficiency, demonstrating exceptional performance due to their high efficiency and improved charge transfer capabilities. 261,262 Bismuth oxyhalides (BiOX, X = Cl, Br, I) present good photoactivity due to their [Bi₂O₂]²⁺ layered structure flanked by the halogen atom, which generates an electrostatic region that inhibits exciton pair recombination.^{263,264} In 2016, Zhu et al. synthesized a hybrid photocatalyst by compounding two-dimensional BiOBr with the MOF to form BiOBr/NH2-MIL-125 aimed at RhB degradation.169 The suitable band gap value and the soformed heterojunction facilitated electron transfer from BiOBr to Ti⁴⁺, enhancing the degradation activity of the composite compared with the individual components.

Direct doping of Ag semiconductors is a straightforward procedure to adjust MILs' band structures. For instance, coating NH2-MIL-125 with silver phosphate nanoparticles to form a well-suited heterogeneous interface of Ag₃PO₄@MIL reduced the $E_{\rm g}$ from 2.51 eV to 2.39 eV, leading to a 39-fold enhancement in photoactivity for MB degradation compared with TiO2.171 Similarly, doping NH2-MIL-125 with Ag3VO4 and Ag_2WO_4 showed that Ag_3VO_4 , with an E_g reduction from 2.65 eV to 2.27 eV, is more effective for optimizing NH2-MIL-125's

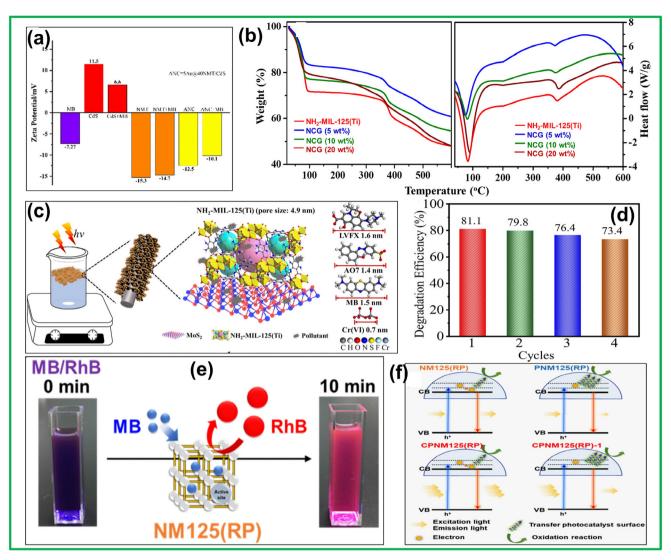


Fig. 7 (a) Zeta potential of MB, CdS, CdS + MB, NMT, NMT + MB, 5Au@40NMT/CdS and 5Au@40NMT/CdS + MB. Reprinted with permission from ref. 166. Copyright 2022 Elsevier B.V. (b) TGA of (left) NH₂-MIL-125(Ti) (right) NH₂-MIL-125(Ti)—CdS/Graphene (20 wt%). Reprinted with permission from ref. 167. Copyright 2021 Elsevier Inc. (c) Photocatalytic reaction device and adsorption process of CF/MoS₂/NH₂-MIL-125(Ti) cloth. (d) Cycle run test in degrading LVFX. Reprinted with permission from ref. 168. Copyright 2021 Elsevier Inc. (e) A schematic representation of NM125(RP)'s selective adsorption capability toward the mixed MB/RhB solution and (f) the proposed mechanism based on the PL results over different samples. Reprinted with permission from ref. 174. Copyright 2023 Elsevier B.V.

band structure. ¹⁷² In 2015 Abdelhameed and colleagues improved the Ti-MOF's photocatalytic efficiency and chemical permanency by integrating it with Cr(III) and Ag NPs. Initially, the NH₂-groups in the MOF underwent a treatment with acetylacetone. ¹⁷¹ When exposed to visible light, both Cr-MIL-125-AC and Ag-MIL-125-AC were able to nearly fully degrade methylene blue (MB). Cr(III) served as an h⁺ acceptor, while Ag NPs functioned as an e⁻ acceptor, thereby enhancing the exciton separation and prolonging their period of recombination. Ag-MIL-125-AC showcased effective and constant photocatalytic performance in degrading MB, maintaining a steady photodegradation rate constant through five consecutive reaction cycles. Ag/AgBr nanoparticles generate charge carriers under visible light owing to AgBr's narrow E_g of 2.6 eV and Ag NPs'

SPR effect.²⁶⁵ Typically, AgBr forms by adsorbing Ag⁺ ions or Br sources onto semiconductors and adding the complementary ion, often resulting in large, aggregated AgBr particles.^{266–268} In 2019, Han *et al.* devised a novel method to produce AgBr NPs and fabricate Ag/AgBr/NH₂-MIL-125 nanocomposites.¹⁷³ Initially, they used amino terephthalic acid and 2,5-dibromo terephthalic acid as linkers to form Br₂-NH₂-MIL-125. Subsequently, AgBr was prepared *via* an *in situ* method by reacting the –Br groups in Br₂-NH₂-MIL-125 with AgNO₃ in ethanol. Finally, Ag/AgBr/NH₂-MIL-125 was fabricated using a photoreduction method, resulting in small, uniformly sized Ag/AgBr nanoparticles (about 4 nm) as a result of the even dispersal of Br in Br₂-NH₂-MIL-125. The prepared hetero composites demonstrated significantly improved light-

harvesting ability and higher photocatalytic efficiency for the removal of MO compared with $Br_2\text{-NH}_2\text{-MIL-125}$. The photocatalytic rate constant of the nanocomposite was roughly 5.1-fold greater than $Br_2\text{-NH}_2\text{-MIL-125}$. This enhanced activity was due to the SPR effect of the Ag NPs and the creation of heterojunctions between the components. This research presented a novel method for developing Ag/AgX/MOF (X = Cl, Br) photocatalysts, which holds promise for environmental purification

and energy utilization.269 Quantum dots (QDs), which are less than 10 nm in size, represent a cutting-edge category of 0D nanomaterials. In 2019, Wang et al. devised an upfront way to embed CODs onto the MOF substrate via a solvent-deposition method.²⁷⁰ The NH₂-MIL-125, characterized by its high SSA, served as an optimal matrix for the dispersion of CQDs. These anchored CQDs maximized the utilization of visible light and reduced the rate of exciton species recombination within the MOF structure. When exposed to a broad spectrum of light, CQDs/ NH₂-MIL-125 demonstrated superior photocatalytic activity for the degradation of RhB compared with pristine NH2-MIL-125, with the optimal performance observed at 1% COD content. This superior photocatalytic activity was attributed to the effective photogenerated e⁻/h⁺ pair separations and enhanced light energy utilization. In 2023, Huang and colleagues incorporated CQDs into P25/NH2-MIL-125 via a simplistic reflux technique, forming a ternary composite that significantly boosted photocatalytic performance by promoting longer charge carrier lifetime and promoting up-conversion luminescence. 174 This enhancement allowed for the alteration of near-IR light to visible light, thus advancing the light implementation efficacy. Post-synthesis treatment at 120 °C elevated the stability of the samples in aqueous environments. The UV-vis study revealed a notable red shifting owing to the presence of CQDs, effectively narrowing the $E_{\rm g}$. Unlike P25 alone, the CQDs were dispersed across the surface and integrated into the NH2-MIL-125's pore structure, resulting in weakened XRD patterns and altered porous structures. CQDs efficiently transformed longer-wavelength light with λ > 500 nm into shorter-wavelength light with λ < 500 nm, thus generating more photoinduced electrons in both P25 and Ti-MOF. Due to their excellent conductivity, CQDs facilitated the effective exciton pair separation, thereby enhancing the overall photocatalytic efficacy (Fig. 7e and f). Boron nitride quantum dots (BNQDs) efficiently separate electron-hole pairs due to their electrostatic affinity for photoinduced positive charges, aided by oxygen-containing groups that carry negative charges. 271,272 In 2023, a range of BNQDs incorporated into NH₂-MIL-125 (NMTB-x) nanocomposites from 0D-3D structures were synthesized through co-precipitation for photocatalytic RhB degradation. 175 NH2-MIL-125 and NMTB-0.7 exhibited SSA of around 1096.82 and 1265.62 m² g⁻¹, respectively, offering increased active sites for the catalytic reaction. The study focused on the visible light photocatalytic activity of NMTB-x composites for degrading RhB and TC. Free radical scavenging tests reveal that O2 - significantly impacted photocatalytic performance, followed by hydroxyl radicals 'OH and

 $h^{\scriptscriptstyle +}$. The improved degradation was due to effective heterojunctions, efficient $h^{\scriptscriptstyle +}$ quenching by BNQDs, and reduced carrier recombination.

10.8. Organic transformation reactions

Photocatalytic transformations of organic compounds are valued for their high selectivity, yield, and ability to avoid thermal side reactions, all under mild conditions without toxic reagents. Recent advancements have shown significant progress in MOF-catalyzed organic transformations under visible light. Examples include the photocatalytic oxidation of alcohols,²⁷³ the photocatalytic α-alkylation of aldehydes,²⁷⁴ and aza-Henry reactions.²⁷⁵ NH₂-MIL-125, Ti-based MOF, has emerged as a versatile heterogeneous catalyst in a number of these transformation reactions. Table 2 lists the organic transformation reaction performances of different NH₂-MIL-125(Ti)-based multifunctional materials.

Zhao and colleagues synthesized NH2-MIL-125/TiO2, which displayed outstanding efficiency in the selective oxidation of cyclohexane.276 This achievement was attributed to its excellent light-harvesting capabilities and abundant catalytic sites. This study investigated not only the rate of reaction and possible yield outcomes but also real-world applications in selective oxidation. Cyclohexane oxidation is an important industrial process, but traditional methods often include high temperature and excessive waste generation. Photocatalytic oxidation may offer a cleaner and more sustainable alternative. In a separate study, Gao et al. developed an innovative CdS/NH2-MIL-125@TiO2 composite by coating NH2-MIL-125 with a super thin layer of TiO2 and embedding CdS within the structure.277 This TiO2 layer served to shield the photocatalysts from nitrate and nitrite corrosion and poisoning by photocatalytically removing NO, thus enhancing the composite's durability and stability. The remarkable optochemical characteristics of the nanocomposite resulted from dual excitation pathways, strong interfacial connectivity, and the well-aligned band potentials of NH₂-MIL-125 and CdS quantum dots. Photoexcited e⁻s were captured by O₂, forming superoxide radicals, which subsequently oxidized NO into HNO2 and HNO3 (Fig. 8a). Each component of the catalyst played a distinct role: the high surface area of the MOF enhanced reactant adsorption, CdS QDs enabled visible light responsiveness, and TiO2 prevented e⁻/h⁺ recombination by acting as a buffer.

Li et al. investigated the photocatalytic efficiency towards aerobic oxidation of amines, converting them to imines in an $\rm O_2$ environment (as an oxidant). Their findings revealed that benzylamine credibly transformed into N-benzylidene benzylamine with the MOF catalyst, achieving a 73% conversion rate and 86% selectivity in CH₃CN after 12 hours of light exposure. The structural integrity of the parent MOF was maintained, and its transformation efficacy remained fairly similar even after three cycles, demonstrating its durability and reusability. Amines with e-donating groups exhibited increased conversion rates (75–92%) and improved selectivity (90–93%) concerning benzylamine. The mechanism for photocatalytic aerobic oxidation involved e $^-$ movement following the LMCT

Fable 2 List of NH2-MIL-125(Ti)-based multifunctional materials towards different organic transformation reactions

Sl. no.	Photocatalytic application	Material	Band React gap (eV) type	Reaction type	Substrate	Product	Photocatalytic efficiency	Light source	Ref.
1	Organic transformation	$\mathrm{NH}_2\text{-MIL-}125(\mathrm{Ti})/\mathrm{TiO}_2$	2.76	Oxidation	Oxidation Cyclohexane	Cyclohexanol and	23% conversion, five h	300 W xenon lamp ($\lambda > 420 \text{ nm}$)	276
2	reaction	${\rm CdS/NH_2\text{-}MIL\text{-}125@TiO}_2$	2.50	Oxidation	ON	HNO ₂ and HNO ₃	48.4% conversion,	300 W xenon lamp with UV	277
3		MR-MIL-125(Ti)	1.93	Oxidation	Benzyl alcohol	Benzaldehyde	IIIII c	cut inter (4 > 420 init) 150 W xenon lamp with	126
4		NH_2 -MIL-125@TAPB-PDA-3 2.61	2.61	Oxidation	Benzyl alcohol	Benzaldehyde	94.7% conversion,	300 W xenon lamp $(\lambda > 420 \text{ nm})$	278
5		Pt/Ti-MOF-NH2	1	Reduction	Nitrobenzene	Aniline	11 00	500 W Xe lamp ($\lambda > 420 \text{ nm}$)	279
9		Ni-doped NH ₂ -MIL-125(Ti)	2.5	Oxidation	<i>p</i> -Methyl benzyl	p-Methyl benzaldehvde	43.2% conversion,	300 W xenon lamp with UV cut-off filter (3 > 420 nm)	280
^		$\mathrm{NH}_2\text{-MIL-}125(\mathrm{Ti})$	2.4	Oxidation	Benzylamine	N-Benzylidene benzylamine	73% conversion,	300 W xenon lamp with UV cut-off filter (2 > 420 nm)	281
8		$\mathrm{SMIL}\text{-}\mathrm{NH}_2$	2.4	Oxidation	Acetophenone oxime	Acetophenone	100% conversion,	Blue LED lamps $(4 \times 3 \text{ W})$	282

process, generating Ti³⁺ centers. These highly reactive Ti³⁺ sites interacted with oxygen to produce superoxide radicals and were subsequently reversely oxidized to Ti⁴⁺. Meanwhile, carbon-centered radicals generated *via* photoillumination reacted with superoxide radicals, forming aldehydes. These aldehydes were then nucleophilically attacked by amines, leading to the formation of imines through a dehydration process. The high efficacy of the reaction and the mild conditions under which it runs, therefore avoiding harsh reagents and conditions, underlines the increasing trend towards "green" catalytic processes.

Gascon and colleagues modified NH2-MIL-125 MOF with a dye-like entity (methyl red), creating MR-MIL-125(Ti). 126 This new material demonstrated enhanced photocatalytic productivity in oxidizing the benzyl alcohol, attributed to its significantly improved light absorption than the original MOF. Likewise, novel NH2-MIL-125 MOFs modified with aromatic heterocycles were developed through a Schiff base reaction to develop the π -delocalized system.⁵¹ 2-Pyridine carboxaldehyde, 3-pyridine carboxaldehyde, 4-pyridine carboxaldehyde, 3-thiophene carboxaldehyde, and 2-quinoline carboxaldehyde were imbedded onto the MOF without affecting its parent framework. Out of these, THI- and OUImodified MOFs demonstrated superior photocatalytic performance due to the enhanced conjugated system, which promoted a longer separation period and transference of the exciton pairs.

Zhao *et al.* developed a range of NH₂-MIL-125@TAPB-PDA nanocomposites using an innovative seed growth technique to coat porous and crystalline COF onto the titanium MOF.²⁷⁸ The COF width was adjustable by varying the concentration of COF raw materials. Notably, NH₂-MIL-125@TAPB-PDA-3, with a 20 nm-thick COF shell, achieved the maximum benzyl alcohol conversion rate of 94.7%, outperforming single NH₂-MIL-125 and COF by 15.5 and 12.5 times, respectively. This significant enhancement in its performance was ascribed to the superior exciton pair transference between the MOF and the COF over covalent bondings. Radical trapping experiments and ESR confirmed the presence of superoxide radical anions and holes, which are crucial for the aerobic oxidation reaction.

The aromatic nitro compounds to amines reduction is crucial due to the value of aromatic amines as intermediates in fine chemical synthesis. Matsuoka and co-workers achieved effective photoreduction of nitrobenzene using Pt/NH₂-MIL-125 under the illumination of visible light.²⁷⁹ The initial conversion rates of nitrobenzene for Pt/Ti-MOF-NH₂ and Ti-MOF-NH₂ were found to be 3.3 and 2.3 mmol h⁻¹, respectively. The enhanced photoactivity of the composite indicated that the Pt NPs acted as an effective co-catalyst. The e⁻s were primarily transported to the actively participating Ti-O cluster and subsequently to the Pt (co-catalyst), reducing the nitrobenzene, while TEOA served as an e⁻ donor (sacrificial agent), being oxidized over the linker of ATA.

Zhu and colleagues examined the capabilities of Ni-doped Ti-based MOF, NH₂-MIL-125, towards aromatic alcohols selectively oxidized to aldehydes, utilizing O₂ as an oxidizing agent

·O₂·OH (a)(c) V vs NHF CB: conduction band VB : valence band CSV: cathodic stripping voltammetry NH₂-MIL-125 oxidized manganese product h+ h+ (d) h+ h+ NH₂-MIL-125 ~2.64 V +2.95 CdS QDs h+ h+ ~2.28 V Amorphous TiO₂ 3.25 V NH₂-MIL-125(Ti)/TiO GOX ho (b) (e) (a.u.) Relative Intensity $^{5}D_{0} \rightarrow ^{7}F_{2} Eu^{3+}$

Fig. 8 (a) Possible mechanism of photocatalytic oxidation of NO over CdS/NH₂-MIL-125@TiO₂ catalysts. Reprinted with permission from ref. 277. Copyright 2018 Elsevier B.V. (b) Proposed mechanism of the photocatalytic amines oxidation over NH₂-MIL-125(Ti). Reprinted with permission from ref. 281. Copyright 2014 Elsevier B.V. (c) Schematic illustration for the photooxidation of Mn²⁺ at NH₂-MIL-125(Ti)/CPE and its detection by CSV. Reprinted with permission from ref. 283. Copyright 2014 Elsevier B.V. (d) The electron-transfer mechanism in $GO_x/CS/NH_2-MIL-125(Ti)/TiO_2$ biosensor. Reprinted with permission from ref. 284. Copyright 2017 The Royal Society of Chemistry. (e) $^5D_0 \rightarrow ^7F_2$ transition intensity ratios from the emission spectra of MIL-125-AM-Eu introduced into α-phenethyl alcohol after UV radiation for different times. Reprinted with permission from ref. 285. Copyright 2016 American Chemical Society.

and under visible light illumination. ²⁸⁰ Ni-doped NH₂-MIL-125 demonstrated significantly improved photocatalytic behavior compared with NH₂-MIL-125, with *p*-methyl benzyl alcohol conversions of 43.2% and 25.8%, respectively. Ni nanoparticles, averaging 3 nm in diameter, were encapsulated within the pores of the MOF. The increased performance of Nidoped MOF was credited to its enhanced visible light absorption efficacy, which facilitated the charge transfer efficiency, a result of Ni doping. Furthermore, analyses including N₂ adsorption–desorption, PXRD, and reusability experiments indicated that the structural integrity and photocatalytic performance of the MOF composite remained nearly unchanged after three continuous cycles, demonstrating its stability.

Morsali *et al.* synthesized a surfactant-aided variant of NH_2 -MIL-125, named SMIL- NH_2 , using the solvothermal technique

with Pluronic P123 – a structure-directing entity. ²⁸² They evaluated the performance of photocatalyst SMIL-NH₂ in deoximation reactions. In this efficient photosystem, SMIL-NH₂ facilitated the regeneration of carbonyl complexes from several oximes at twice the rate achieved by NH₂-MIL-125 (Scheme 1). Upon achieving a 100% yield in the photocatalytic deoximation of oximes using SMIL-NH₂, the conversion time for acetophenone oxime to its corresponding ketone was about 10 hours. For acetophenone oxime derivatives with e⁻ donating methyl and methoxy groups, the reaction times were significantly reduced. Moreover, SMIL-NH₂ exhibited notable structural stability and cycling stability while maintaining its photocatalytic activity intact. However, long-term stability and metal-deposited composites. Optimizing heterojunction structures is

5 min 10 min 15 min 20 min 25 min 30 min

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crucial for efficient electron-hole separation, requiring precise control of coating thickness and interfaces.

10.9. Photocatalytic sensors

Photocatalytic sensors, which integrate irradiation of photon energy with electrochemical detection, have gained considerable interest for their capacity to combine the benefits of optical techniques with those of electrochemical sensors. Tibased MOFs, renowned for their excellent photocatalytic and redox characteristics, are particularly promising for photoelectrochemical sensors. Table 3 lists the different types of modified NH₂-MIL-125(Ti) multifunctional nanomaterials and their performances towards photocatalytic sensing. In one study, Hu and colleagues enhanced a carbon paste electrode by incorporating NH₂-MIL-125 as the working electrode. By the combination of visible light and CSV methods regarding the oxidized Mn materials, they detected Mn²⁺ quantitatively (Fig. 8c).²⁸³ Upon exposure to light, exciton pairs transferred to the NH₂-MIL-125 surface and reacted with Mn²⁺ species in the mixture solution, facilitating the generation and preconcentration of MnO₂. The detection of this product was achieved using the CSV process, which is responsible for quantifying the oxidation products' stripping peak current. The peak currents exhibited a linear increase with the concentration of Mn²⁺, demonstrating a strong linear correlation across the range of 1.0×10^{-8} to 1.0×10^{-5} M. This work was particularly remarkable with successful application on real-world samples, including the detection of Mn²⁺ in tea, one of the challenging aspects for most sensing systems that go well in lab conditions alone. The same group also fabricated a photoelectrochemical sensor ground on an NH₂-MIL-125/TiO₂ nanocomposite for determining the herbicide, clethodim. They coated the as-prepared composite on a glassy carbon electrode that served as the working electrode.²⁸⁶ In this case, during the sensing process, the interaction between photogenerated h⁺ on the nanocomposite reacted with water to produce 'OH radicals. The herbicide then quickly interacted with these 'OH radicals, which improved charge-hole separation efficiency and increased the photocurrent, providing a reliable and sensitive detection method. The photocurrent determined clethodim concentration, displaying a strong linear correlation and a lower detection margin of 10 nmol L⁻¹. This work showcases the significance of nanocomposite materials for optimizing the performances of photocatalytic sensors by enhancing charge transfer efficiencies and sensor response times.

Subsequently, a more intricate yet sensitive photoelectrochemical sensor was created using a GO_x/chitosan/NH₂-MIL-125(Ti)/TiO2 nanocomposite. This sensor was designed to detect the pesticide acetochlor present in agricultural produce (Fig. 8d).²⁸⁴ This multi-functional sensor, which integrates the interaction between photogenerated radicals with the pesticide, illustrates the flexibility of photocatalytic sensors in terms of integrating biological elements such as enzymes for further enhancement of selectivity. In this setup, CS was used as a matrix to immobilize GOx, facilitating the conversion of glucose to H_2O_2 . The H_2O_2 acted as an e-donor, scavenging h⁺

Table 3 List of NH₂-MIL-125(Ti)-based multifunctional materials towards photocatalytic sensing applications

Sl. no.	Photocatalytic I. no. application Material	Material	Band gap (eV)	Target pollutant/ analyte	Light source	Sensor configuration	Detection limit	Detection limit Stability (cycles) Ref.	Ref.
7 7	Sensors	$\mathrm{NH_2\text{-}MIL\text{-}}125$ $\mathrm{NH_2\text{-}MIL\text{-}}125/\mathrm{Ti}\mathrm{O}_2$	2.25 2.25	Mn ²⁺ Clethodim	250 W halogen lamp	Modified CPE Modified GCE	4.0 nmol L^{-1} 10 nmol L^{-1}	>100 cycles >100 cycles	283 286
3		$\mathrm{NH_2\text{-}MIL\text{-}}125/\mathrm{TiO_2}$	2.5	Acetochlor	(4 = 340-850 nm) 250 W halogen lamp	$GO_x/CS/NH_2-MIL-125(Ti)/TiO_T$	$0.003 \mathrm{\ nmol\ L^{-1}}$	I	284
4		MIL-125(Ti)-AM-Eu	I	α-Phenethyl alcohol	(A > 400 mm) UV light	Photocatalyst & fluorescence	I	5 cycles	285

in the VB of the nanocomposite, thereby enhancing charge separation efficiency and increasing the photocurrent. However, in the presence of acetochlor, the photocurrent significantly lessened because acetochlor inhibited GO_x enzyme activity. The concentration of acetochlor was quantified by measuring the fraction of inhibition of GO_x activity. This analytical process, which relies on the photocurrent being contrary to the acetochlor concentration, was effectively used to detect acetochlor in fresh produce samples. The vision of this study appears to be compelling because it explains ways photocatalytic materials can be used with bio-recognition elements to tackle particular analytes, an important advancement in developing highly selective sensors. However, it is more likely that enzyme-based sensors will be unstable, especially under varying external environment conditions. Further optimization of the composite and enzyme-immobilization process may reduce this issue.

Yan and colleagues created an Eu-modified NH2-MIL-125, MIL-125-AM-Eu, as a fluorescence sensor (turn-off) for detecting α -phenethyl alcohol.²⁸⁵ When exposed to UV light, the sensor exhibited fluorescence from the e⁻ transition of ${}^5D_0 \rightarrow$ ⁷F₂ 4f in Eu³⁺ attached to MIL-125-AM. The incorporation of lanthanide functionalization introduces a new aspect to photocatalytic sensing. Rare-earth elements like Eu provide improved sensitivity thanks to their distinct electronic properties and their capacity for sharp, well-defined electronic transitions. This fluorescence intensity gradually decreased and was fully snuffed in about 20 minutes (Fig. 8e). α-Phenethyl alcohol became trapped in the MIL-125-AM-Eu channels; as a result, α-phenethyl alcohol oxidized to acetophenone, facilitated by the photoactive Ti(IV). This reaction led to fluorescence quenching due to the formation of carbonyl groups. The consistent luminescent emission intensity, photocatalytic transformation, and Ti to Eu proportion with five reaction cycles demonstrated the material's reusability for detecting α-phenethyl alcohol. The MIL-125-AM-Eu's excellent performance validated its effectiveness as a turn-off fluorescence sensor towards α -phenethyl alcohol. Moreover, researchers have further discovered that MIL-125-AM-Eu can also act as a ratiometric fluorescence sensor for anion recognition and detection. Enhancing sensor sensitivity and selectivity for accurate and reliable analyte detection is a key focus. Challenges include ensuring stability under operational conditions and optimizing charge separation and photocurrent response. Ensuring long-term stability and reusability, expanding detectable substances, and implementing sensors in real-world applications are crucial for advancing the field.

11. Conclusion

Even though Ti-based MOF materials constitute only a compact subset of MOF descendants, significant progression in their structural design and fabrication has been made in recent years. In the case of metal-modified MOFs, like Zr, Cu

and Fe-MOFs, Ti-based MOFs offer greater structural and topological divergence in titanium cores owing to their varying nuclearity. However, Ti-MOF structures remain relatively scarce compared with di- and trivalent metal MOFs and Zr-MOFs, primarily because of the increased volatility and hydrophilicity of ionic Ti forerunners. Despite years of research leading to a wealth of Ti-cluster coordination chemistry, the field of Ti-MOFs only saw a breakthrough in recent years with the finding of MIL-125, a significant juncture. However, a higher $E_{\rm g}$ (3.6 eV) of MIL-125 curbs its activity in the UV region. To address this, NH₂ (organic chromophore) was introduced to extend the light absorption range into the visible light region. The welldefined but tunable structure of NH2-MIL-125 has attracted intensive research due to its robust architecture and intriguing photoresponsive properties. This makes it a promising material for assembling multifunctional photocatalysts and an ideal platform for mechanistic studies of photocatalytic processes. The review elucidates the critical aspects, such as the nature of the Ti precursor, coordination environment, synthesis procedures, and charge transfer mechanisms, providing a comprehensive understanding of their photocatalytic behavior. Various strategies were also discussed to enhance lightharvesting ability and boost exciton pair split for improved photocatalytic efficacy. These strategies aim to optimize the optical properties, enhance the electron transfer rate, and accelerate surface catalytic reactions in the NH₂-MIL-125 matrix. The broad realm of photocatalytic applications for NH₂-MIL-125 nanomaterials, including hydrogen and oxygen evolution reactions, CO2 and N2 reduction, photocatalytic sensors, and organic transformations, highlights their versatility and potential impact.

12. Challenges and future prospects

The aminated MIL-125(Ti) has been exemplified as a photocatalyst for numerous photocatalytic applications. The multinuclear titanium centers, akin to those in TiO₂, endow the MOF with significant photocatalytic potential. Additionally, the titanium-based MOF features outstanding porosity and structural sturdiness, variable valence states of Ti(IV/III), and lower CB concerning Zr-MOFs, which promotes a durable LMCT with photoirradiation, making it an appropriate selection for photocatalytic applications. Despite some eye-catching and notable results in NH₂-MIL-125-related photocatalysis, certain challenges are still faced by the MOF concerning the structure, morphology, and band gap.

i. Ti precursors spontaneously hydrolyze with water, whether present as solvent residue, intentionally added, or generated *in situ* from reactions. This typically results in the development of huge multinuclear SBUs and amorphous TiO₂, partially blocking the cavities of the newly developed MOFs. Enhancing the complex potential of ligands can lead to compact Ti-oxo clusters or surprisingly remote Ti octahedra while suppressing the competitive side formation of TiO₂. There have been reports of only a few systems, mainly invol-

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ving common organic ligands, but there remains considerable room for exploring new NH2-MIL-125-based structures.

ii. Functionalizing MIL-125 analogs has shown promise but is limited. Lowering the E_g below 1.3 eV is difficult due to the large Ti₈-oxo clusters, which form local semiconductive regions and are minimally affected by ligand modifications. Expanding the absorption band may not enhance photocatalytic activity, as low-energy photons insufficient for photoreactions are also absorbed. Additionally, added groups can block MOF apertures and impede mass transport.

iii. Synthesis remains the primary challenge in NH2-MIL-125 research, primarily due to the greatly unpredictable Ti⁴⁺ cation polycondensation, leading to unrestricted products. Synthetic parameters, such as solvent composition, time, and reaction temperature, besides the surfactant and modulator addition, significantly affect the end products' design, porosity, and crystalline morphology.

iv. The preparation and usage of NH2-MIL-125 suffer from difficulties in the handling of large Ti-oxo-carboxylate clusters, while NH2-MIL-125 has a lower surface area compared with MIL-125 due to partial pore blockage by the free ligand, and both reveal a reversible decrease in crystallinity upon highpressure treatment.

v. The properties of NH2-MIL-125 with a mouldable framework are capable of being regulated with the introduction of counterfeiting groups to the linkers or substituting/doping the nodes with the metal ions or the clusters. However, reports of the above-modified MILs are scantily available.

vi. The coordinated NH₂-MIL-125 structure allows the integration of multifunctional modules into an individual framework towards collaborative catalysis, facilitating the construction of diverse MOF-based nano photocatalysts having multiple functions. Yet, the most complex NH2-MIL-125-based

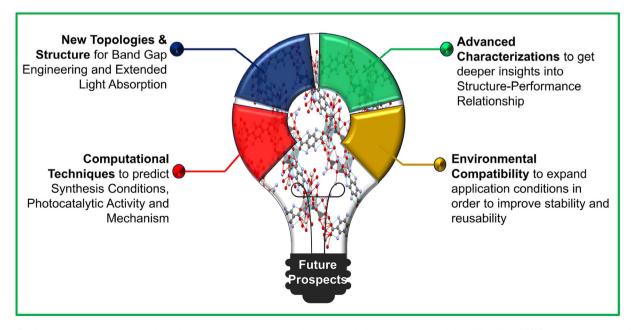
materials show an unclear structure, significantly hindering the study of the relationship concerning structure, property, and performance.

vii. Exploring NH₂-MIL-125-based photocatalysts towards artificial photosynthetic overall processes is a challenging yet significant endeavor, encompassing water splitting, hydrogen peroxide production, artificial carbon dioxide conversion, and nitrogen reduction.

viii. NH2-MIL-125-based photocatalysts generally exhibit a moderate or low photocatalytic efficiency, necessitating significant improvements for practical applications. The interdisciplinary integration of MOF research and computational techniques - such as machine learning, big data, artificial intelligence, and the Internet of Things - holds the potential to address this challenge.

ix. Presently reported, light harnessing NH2-MIL-125-based catalysts are budding towards CO2 reduction and H2 generation, relying heavily on sacrificial agents that are neither cost nor environment-friendly. Besides, research on NH2-MIL-125based photocatalysts remains at the laboratory scale. Catalytic efficiency depends not only on the photocatalyst's properties but also on environmental factors. In real-world conditions, these photocatalysts face challenges such as ion interference, temperature fluctuations, and pH variations (Scheme 3).

The capabilities of NH₂-MIL-125(Ti) are crucial owing to its potential to revolutionize several critical areas, including advanced materials science. Enhancing its photocatalytic efficiency means contributing significantly to renewable energy solutions and pollution control by enabling more effective solar-driven processes. Improvements in gas storage and separation capabilities help to form better energy storage systems and more efficient capture of greenhouse gases, addressing urgent environmental concerns. Furthermore,



Scheme 3 Future prospects towards the advancements and enhanced photocatalytic performance of the NH2-MIL-125(Ti)-based nanomaterials.

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advancements in sensor technologies using NH2-MIL-125(Ti) could result in useful and selective devices for detecting pollutants and hazardous substances, promoting public health and safety. Therefore, research in this space is still in its initial phases, providing enough scope for further expansion and optimization of NH₂-MIL-125(Ti), vital for advancing technologies supporting a more sustainable and healthier future. Hence.

- i. Despite the challenges, new topology and structure construction is the need of the hour for the bandgap engineering of NH2-MIL-125-based photocatalysts, to underscore their full potential. The bandgap of the MOF can be engineered by functionalizing ligands and developing new structures.
- ii. Currently, the direct solvothermal method is the most customary synthesis procedure for producing this MOF. However, controlling the product structure is hugely empirical and particularly difficult, even though not impossible, while exploring new systems. Further work is needed on optimizing synthesis for better surface area and stability, using different BDC derivatives to modulate photocatalytic properties, and exploring its mechanical stability and structural behaviors under various stimuli to extend the usage of MOFs. Also, it is possible to impromptu assemble SBUs through the coordination of covalent bonds and frameworks in a one-pot reaction. Titanium is particularly well suited for this approach owing to the strength and abundance of Ti-oxo-clusters. The combination of a one-pot reaction strategy with other techniques eventually enables the target-specific synthesis of NH₂-MIL-125 having pre-formed topology and chemical functionality.

iii. Investigation of the morphology and the interface properties of the various components forming a heterojunction with MOF and alteration of its organic ligands/metal nodes via thorough research can help in developing multiple NH2-MIL-125 structures to improve its photocatalytic efficiency, stability, and reusability. Besides, catalytic ability can be further enhanced by incorporating different hugely active species, such as nanoparticles of metal and graphene. Additionally, new NH2-MIL-125-based composites should be explored, having explicit textural structures, low E_g , and high energy transfer efficiency.

- iv. While expanding the applications of NH2-MIL-125 is crucial, an ultimate goal for this family of materials is the efficient employment of solar energy through its distinct photoactive moieties, spanning the absorption of solar energy into the visible and NIR regions. Also, the E_g engineering of the MOF can be brought about by functionalizing ligands to develop innovative structures.
- v. Additionally, using NH₂-MIL-125(Ti)-based photocatalysts for photocatalytic oxidative desulfurization could be a groundbreaking approach for addressing environmental issues.
- vi. Advanced computational techniques could efficiently identify optimal synthetic conditions for tuning the morphology or the MOF for specific photoassisted catalytic reactions. For instance, artificial intelligence involving machine learning aims to design models based on previous literature. A

standard machine learning model, capable of predicting electronic E_g and photocatalytic activities and guiding the engineering of morphology, could provide new perceptions for the well-reasoned blueprint of NH₂-MIL-125-based photocatalysts with enhanced catalytic performance utilizing solar energy.

vii. To address the issue of the need for sacrificial agents and to eliminate them, a combination of oxidation side reactions, such as oxygen evolution reaction (OER), with reduction reactions is crucial. Additionally, the use of single atoms as cocatalysts can assist the evolution of hydrogen in place of precious metals.

viii. The mechanistic pathway of photocatalytic reactions over NH2-MIL-125-based photocatalysts remains unclear. Advanced characterization procedures such as in situ analysis, flash-photolysis without electrodes, transient absorption techniques, time-dependent microwave conductivity, and isotopic methods, along with DFT (Density Functional Theory) calculations, can be feasibly employed to gain deeper insights into the performance and structure relationship. This understanding is crucial for the well-reasoned design of MOFs (Ti-based) with enhanced photocatalytic efficiency.

ix. Future research must also focus on the interrelation between NH2-MIL-125-based photocatalysts and their surroundings, expanding the application horizon of the MOF. Additionally, stability and recyclability are pivotal for the practical implementation of these materials. Therefore, further research is required to investigate the causes and prerequisites of the structural disintegration of light-harnessing NH₂-MIL-125-based catalysts.

Although still in its early stage, the rapid developments in the realm of NH₂-MIL-125 as a photocatalyst envision a bright future for the above nanomaterial. We believe that with the help of rationally designing MOFs' structure, we likely look forward to enhanced performance pertaining to photocatalytic solar-to-energy conversion. Through our review, our wish and aim are to inspire researchers to explore novel MOF structures based on NH₂-MIL-125 towards fresh applications carrying conceptual innovativeness in the existing and upcoming intriguing fields.

Author contributions

Priyanka Priyadarshini: conceptualization, writing-original draft, visualization. Anshumika Mishra: writing-original draft, review, and editing. Susanginee Nayak: writing-original draft. Kulamani Parida: conceptualization, visualization, supervision.

Data availability

The submitted review article synthesizes and discusses data that have already been published and publicly available from the sources cited in the manuscript. No new data were generated during the preparation timeframe of this review.

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Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The authors express their profound gratitude toward Siksha 'O' Anusandhan (deemed to be University) for giving all the necessary facilities and financial support to carry out this immense research work.

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