

Cite this: *Chem. Sci.*, 2024, 15, 2914

All publication charges for this article have been paid for by the Royal Society of Chemistry

Received 13th October 2023
Accepted 15th December 2023

DOI: 10.1039/d3sc05447k

rsc.li/chemical-science

Catalytic ammonia synthesis on HY-zeolite-supported angstrom-size molybdenum cluster†

Satoshi Kamiguchi,^{ID} *^{ab} Kiyotaka Asakura,^{ID} ^c Tamaki Shibayama,^{ID} ^d
Tomoko Yokaichiya,^{ID} ^e Tatsushi Ikeda,^{ID} ^e Akira Nakayama,^{ID} *^e
Ken-ichi Shimizu,^{ID} ^c and Zhaomin Hou,^{ID} ^{ab}

The development of new catalysts with high N₂ activation ability is an effective approach for low-temperature ammonia synthesis. Herein, we report a novel angstrom-size molybdenum metal cluster catalyst for efficient ammonia synthesis. This catalyst is prepared by the impregnation of a molybdenum halide cluster complex with an octahedral Mo₆ metal core on HY zeolite, followed by the removal of all the halide ligands by activation with hydrogen. In this activation, the size of the Mo₆ cluster (ca. 7 Å) is almost retained. The resulting angstrom-size cluster shows catalytic activity for ammonia synthesis from N₂ and H₂, and the reaction proceeds continuously even at 200 °C under 5.0 MPa. DFT calculations suggest that N≡N bond cleavage is promoted by the cooperation of the multiple molybdenum sites.

Introduction

Ammonia production from atmospherically abundant dinitrogen (N₂) is an essential chemical process for human beings, because ammonia is a feedstock for globally used fertilizer and a wide variety of nitrogen-containing chemicals.¹ Recently, ammonia has also been expected to be used in hydrogen (H₂) storage owing to its high hydrogen content ratio (17.6 wt%)² and as a fuel that does not emit carbon dioxide during combustion.³ The worldwide production of ammonia is increasing year by year and reached 0.18 Gton in 2022.⁴ Industrially, ammonia is produced from N₂ and H₂ via the Haber–Bosch process using iron (Fe)-based catalysts under high pressure (10–30 MPa) and high temperature (400–500 °C) conditions.⁵ The reaction conditions of industrial ruthenium (Ru)-based catalysts developed later are still harsh (<10 MPa and 325–450 °C).⁶ Making these reaction conditions milder (e.g. <5 MPa and <200 °C) is in high demand to reduce the high energy consumption.⁷ In ammonia synthesis on these industrial catalysts, the cleavage of the chemically inert N≡N triple bond has the highest energy

barrier,⁸ and acceleration of this triple bond cleavage facilitates efficient ammonia synthesis at lower temperatures.^{2c,9} Recently, it has been reported that trinuclear titanium (Ti) and chromium (Cr) hydride cluster complexes achieve N≡N bond cleavage and subsequent hydrogenation of N at ambient temperature and pressure, based on cooperation by multiple metal sites.¹⁰ Biologically, nitrogenase enzymes in certain microbial organisms produce ammonia from N₂ under ambient conditions,¹¹ and the cooperation by multiple metal sites is considered to be responsible for the reaction.¹² These results suggest the potential of clusters for efficient catalytic ammonia synthesis.

Previously, Kamiguchi, one of the authors of this paper, reported that transition-metal cluster compounds with chloride or bromide ligands had catalyzed various reactions since 2002, although there had been no reports on these clusters as catalysts for more than 140 years.¹³ A molecular molybdenum (Mo) chloride cluster with an octahedral metal framework, (H₃O)₂[(Mo₆Cl₈)Cl₆]·6H₂O (**1**), exhibits catalytic activity by partial elimination of halide ligands.¹⁴ In this activation, however, only some of the metallic sites of the Mo₆ cluster can participate in the catalytic reactions. When all the halide ligands are removed by H₂-activation in the anticipation of the participation of more metallic sites, the Mo₆ cluster aggregates to form bulk molybdenum metal.¹⁵ Thus, we expect that the H₂-activation of **1** dispersed on a porous material would form an isolated ultra-small molybdenum metal cluster without aggregation of the Mo₆ cluster core, leading to the cooperation by more molybdenum sites for the N≡N bond cleavage and further efficient ammonia synthesis. Moreover, in comparison with Fe and Ru, Mo is more active in N₂ activation, as deduced from the calculated N₂ dissociative adsorption energy on the metal surface.¹⁶ Several molecular Mo complexes afford

^aAdvanced Catalysis Research Group, RIKEN Center for Sustainable Resource Science, 2-1 Hirosawa, Wako, Saitama, 351-0198, Japan. E-mail: kamigu@riken.jp

^bOrganometallic Chemistry Laboratory, RIKEN Cluster for Pioneering Research, 2-1 Hirosawa, Wako, Saitama, 351-0198, Japan

^cInstitute for Catalysis, Hokkaido University, Sapporo, 001-0021, Japan

^dCenter for Advanced Research of Energy Conversion Materials, Hokkaido University, Sapporo, 060-8628, Japan

^eDepartment of Chemical System Engineering, Graduate School of Engineering, The University of Tokyo, Tokyo, 113-8656, Japan. E-mail: nakayama@chemsys.t.u-tokyo.ac.jp

† Electronic supplementary information (ESI) available. See DOI: <https://doi.org/10.1039/d3sc05447k>

ammonia from N_2 using proton sources and reducing agents at ambient temperature and pressure in a homogeneous system.¹⁷ From these results, it is expected that the ultra-small Mo cluster will show high catalytic activity for Haber–Bosch-like ammonia synthesis from N_2 and H_2 . Herein, we report the preparation of an angstrom-size Mo metal cluster from HY-zeolite-supported **1**. The resulting cluster produced ammonia with high stability, and the reaction proceeded continuously even at 200 °C under 5.0 MPa. The cooperation of the multiple Mo sites promotes the $N\equiv N$ bond cleavage and efficient ammonia synthesis.

Results and discussion

Activation of cluster

As previously reported, when unsupported **1** is heated progressively in flowing hydrogen, **1** is converted to bulk molybdenum metal up to 600 °C, with complete removal of the chloride ligands as hydrogen chloride.¹⁵ When HY-supported **1** was analyzed by H_2 -temperature-programmed reduction (H_2 -TPR), a large reduction peak was observed at 450–600 °C (Fig. S1†), suggesting that **1** on HY also releases chloride ligands as hydrogen chloride up to 600 °C. Thus, supported **1** was heated in flowing hydrogen at 600 °C for 3 h at atmospheric pressure before ammonia synthesis. Elemental analyses showed a decrease in Cl-content from 1.9 to 0 wt% without the loss of Mo-content after the H_2 -treatment (Table S1†), indicating the complete removal of the chloride ligands of **1** on HY by H_2 -activation. HZSM5- and MCM41-supported **1** also showed complete removal of the chloride ligands, as confirmed by elemental analyses (Table S1†).

The change in the local structures of HY-supported **1** by impregnation and H_2 -activation was investigated using the X-ray absorption fine structure (XAFS) technique. The results are summarized in Fig. 1 and Table 1. Cluster **1** impregnated on HY exhibited very similar XAFS (Fig. 1A(b)), X-ray absorption near edge structure (XANES) (Fig. 1A(b), inset), and Fourier transforms of k^3 -weighted extended XAFS (FT-EXAFS) spectra (Fig. 1B(b)) to those of **1** before impregnation (Fig. 1A(a) and B(a)). The fitted parameters of **1** after impregnation were almost the same as those before impregnation (Table 1). Thus, impregnation did not change the molecular structure of **1** or the size of the Mo_6 cluster (*ca.* 7 Å). However, the XAFS (Fig. 1A(c)), XANES (Fig. 1A(c), inset), and FT-EXAFS spectra (Fig. 1B(c)) changed after H_2 -activation. The FT-EXAFS spectrum had two small peaks at 1.5 and 2.5 Å. As shown by curve fitting analysis (Table 1), while the former peak was attributed to Mo–O (oxygen of a silanol of HY), the latter was assigned to the nearest Mo–Mo with a bond length of 2.84 Å. No significant peaks attributed to the next nearest Mo–Mo shell were observed over the longer range (Fig. 1B(c)), indicating that the Mo_6 cluster on HY did not aggregate to form a larger molybdenum particle with long Mo–Mo distances. The coordination number (CN) of the nearest Mo–Mo was 3.6 (Table 1), and this CN value was close to that of **1** before and after impregnation (4.0 and 4.2, respectively), indicating that the average nuclearity of the Mo cluster was almost retained after H_2 -activation. The XAFS (Fig. 1A(e)), XANES (Fig. 1A(e), inset), and FT-EXAFS spectra (Fig. 1B(e)) of Mo foil were quite different from those of **1**/HY after H_2 -activation, which also demonstrates that the impregnation of **1** on HY prevented the aggregation of **1** after activation.

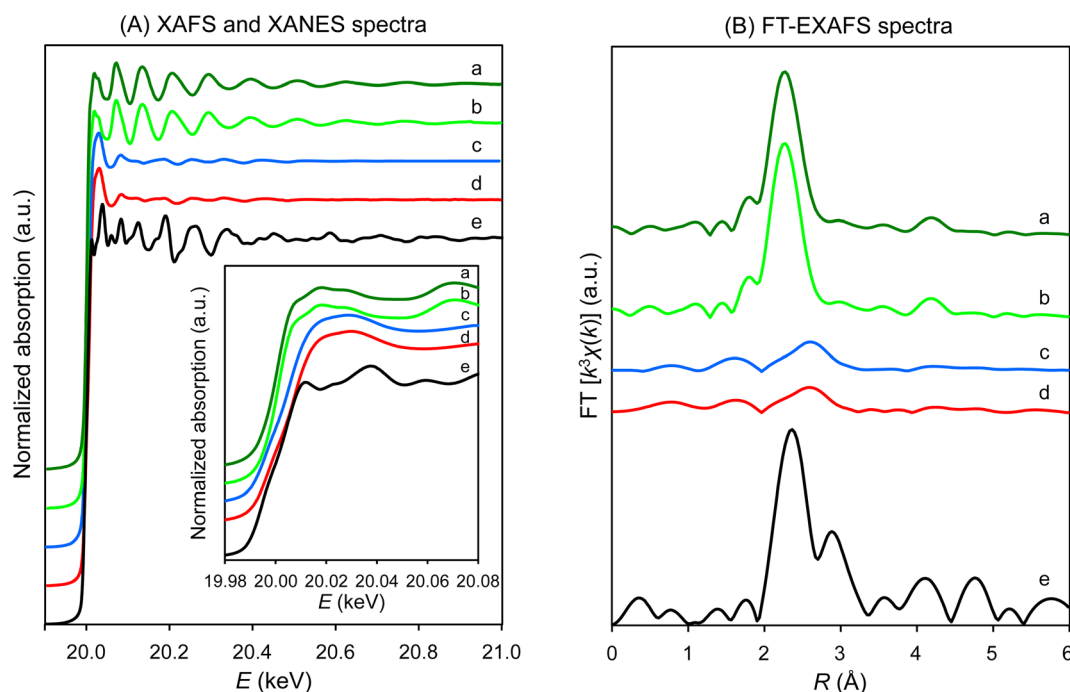


Fig. 1 (A) XAFS spectra of $(H_3O)_2[(Mo_6Cl_8)Cl_6] \cdot 6H_2O$ (**1**). The inset shows the XANES region. (B) FT-EXAFS spectra of **1**: (a) as prepared, (b) **1**/HY after impregnation, (c) **1**/HY after H_2 -activation, and (d) **1**/HY after NH_3 -synthesis. Spectrum of Mo foil from the Spring-8 BENTEN database (<https://doi.org/10.48505/nims.2249>) (e) is also shown.



Table 1 Curve fitting results of Mo K-edge EXAFS data

Sample	Conditions	Shell	CN ^a	R ^a (Å)	$\sigma^a/10^{-2}$ (Å)	R _f ^b (%)
(H ₃ O) ₂ [(Mo ₆ Cl ₈)Cl ₆]·6H ₂ O (1)	As prepared ^c	Mo–Mo	4.0 (fixed)	2.65 (0.01)	2.2 (0.7)	0.14
		Mo–Cl	5.0 (fixed)	2.54 (0.01)	1.3 (1.3)	
(H ₃ O) ₂ [(Mo ₆ Cl ₈)Cl ₆]·6H ₂ O (1)/HY	After impregnation	Mo–Mo	4.2 (0.3)	2.65 (0.01)	2.2 (0.7)	0.12
		Mo–Cl	5.3 (0.2)	2.54 (0.01)	1.3 (1.3)	
	After H ₂ -activation	Mo–Mo	3.6 (0.8)	2.84 (0.01)	7.0 (0.8)	3.2
		Mo–O	1.3 (0.5)	2.04 (0.02)	5.8 (3.2)	
	After NH ₃ -synthesis	Mo–Mo	3.6 (0.7)	2.83 (0.01)	7.7 (0.6)	2.4
		Mo–O	1.0 (0.3)	2.04 (0.02)	3.5 (3.3)	

^a Numbers in parentheses are errors estimated using the Hamilton ratio test with a significance level of 0.317.³⁸ ^b The good fit of the observed and calculated data was also demonstrated by the EXAFS-fitting curves shown in Fig. S13(a)–(d). ^c Sample diluted with boron nitride was analyzed.

The morphologies of the HY-supported **1** after impregnation and H₂-activation were observed using high-magnification Cs-corrected scanning transmission electron microscopy (Cs-STEM). The results are shown in Fig. S2, S3, and S11(a), (b).[†] In the STEM images after impregnation (Fig. S2[†]), each of the white particles corresponds to intact **1**, as the XAFS confirmed the retention of the structure of **1** (see above). The average particle size of 9.7 Å (Fig. S11a[†]) was smaller than the molecular size of **1** including the chloride ligands (*ca.* 11 Å), which is attributed to the lower intensity of Cl than that of Mo, since the clarity of the STEM image is proportional to the square of the atomic weight (36.5 for Cl vs. 96.0 for Mo).¹⁸ Although we tried to obtain a clear atomic image of the Mo clusters, it was impossible because of the fluctuations of the clusters under high energy electron irradiations. The size of the zeolite micropore (7.5 Å) was smaller than that of **1**, and the particles on the edge of the sample projected into the black area (Fig. S2b[†]). These results show that **1** was not embedded in the micropore after impregnation. After H₂-activation, the STEM average particle size decreased from 9.7 to 7.9 Å (Fig. S11b[†]), which was close to the size of the Mo₆ cluster (*ca.* 7 Å), and this decrease is attributed to the removal of chloride ligands by activation. The STEM image shows that, except for some large particles overlapping a stripe, the particles were observed between pore stripes or on the edge of a stripe (Fig. S3a[†]). The pore volume measurements also confirmed the embedding of Mo particles in the micropores after activation (Table S2[†] and description therein). These results indicate that the encapsulation of the metal cluster in the zeolite pore after activation prevented the aggregation of the cluster and retained its average particle size by the interaction of the cluster with silanols in the pore.¹⁹ In spite of some reports on supported molybdenum nitride and carbide clusters,^{20,21} there have been no reports on supported molybdenum metal clusters.

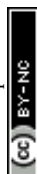
We also analyzed the HZSM5- and MCM41-supported clusters. The XAFS and STEM of **1**/HZSM5 after H₂-activation showed the formation of body-centered cubic (bcc) structured large Mo particles (Fig. S12Bc[†]) with an average particle size of 37 Å (Fig. S11e[†]). All the particle sizes (>10 Å) were larger than the zeolite pore size (*ca.* 5.5 Å). These results indicate that the Mo₆ metal cluster formed by activation was not encapsulated in the small pore but was aggregated to form large bcc-Mo metal

particles outside the pore. Conversely, the MCM41-supported cluster after activation exhibited a smaller CN of Mo–Mo (3.0) in the XAFS (Table S3[†]) and a smaller STEM average particle size (5.5 Å) (Fig. S11h[†]) than the HY-supported cluster. All the particles were smaller (<20 Å) (Fig. S11h[†]) than the MCM41 pore size (*ca.* 24 Å) and were observed on the edge of a stripe (Fig. S9[†]). These results suggest that the Mo₆ metal cluster was converted to a smaller cluster inside the large mesopore. It is reported that H₂-activation of a silica-supported dinuclear Mo complex and its ligand elimination causes coordination of silanol-oxygen atoms to the Mo atoms and cleavage of the Mo–Mo bond.²² The reduction in nuclearity of the Mo₆ metal cluster in the MCM41 mesopore can be explained in the same way. In contrast, the micropore of HY just fits the Mo₆ metal cluster, and therefore, the cluster is embedded without decomposition even when silanol-oxygen atoms coordinate to the Mo atoms. Thus, the size of the Mo metal cluster after H₂-activation depends on the pore size of the support, and HY with a pore size of *ca.* 7.5 Å is suitable for the retention of the size of the Mo₆ cluster of **1**.²³

Catalytic performance

The H₂-activated clusters on the three supports were applied to ammonia synthesis. After the preparation of the H₂-activated clusters, they were subsequently subjected to the reaction of a mixture of N₂ and H₂ with a flow ratio of 1 : 3 at 400 °C and 1.0 MPa (absolute pressure) without exposure to air. Ammonia was continuously formed for 8 h after the start of the reaction for the three supports (Fig. 2). On all these three supports, the clusters after ammonia synthesis showed very similar elemental analysis data (Table S1[†]), XAFS results (Fig. 1, Table 1, Fig. S12, and Table S3[†]), and STEM data (Fig. S4, S7, S10, S11c, S11f, and S11i[†]) to those after H₂-activation. Thus, the structures of the cluster catalysts were stable during the ammonia synthesis.

Table 2 lists ammonia synthesis rates of various supported Mo catalysts at 400 °C. In comparison with the ammonia synthesis rates of H₂-activated **1**/MCM41 (10.2 mmol g_{Mo}^{−1} h^{−1}, entry 1) and **1**/HZSM5 (14.7 mmol g_{Mo}^{−1} h^{−1}, entry 2), that of activated **1**/HY (20.5 mmol g_{Mo}^{−1} h^{−1}, entry 3) was higher at an absolute pressure of 1.0 MPa. This indicates that the Mo cluster with an averaged structure of Mo₆ was most effective for ammonia synthesis. As Table S4[†] shows, the cluster on HY after



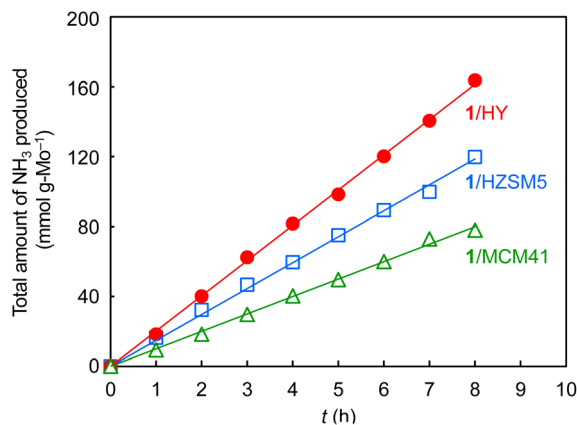


Fig. 2 Catalytic performance of ammonia synthesis at 400 °C and 1.0 MPa (absolute pressure) using $(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (1)/HY, $(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (1)/HZSM5, and $(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (1)/MCM41 as precursors. Catalyst amount, 0.2 g; N_2/H_2 (1/3-mixture) flow rate, 60 mL min^{-1} .

H_2 -activation adsorbed more ammonia ($0.58 \text{ cm}^3 \text{ g}_{\text{cat}}^{-1}$) than those on HZSM5 or MCM41 (0.39 and $0.42 \text{ cm}^3 \text{ g}_{\text{cat}}^{-1}$, respectively), and hence had the largest number of catalytically-active Mo atoms, which could be a cause of the higher ammonia synthesis rate. The smaller number of active sites of the larger-sized Mo cluster on HZSM5 is attributable to the embedding of some Mo atoms inside the large cluster, while that of the smaller-sized cluster on MCM41 is ascribed to the larger ratio of silica-coordinated Mo atoms, as confirmed by the higher CN of Mo–O (2.0) than that for the cluster on HY (1.3) in the XAFS analysis (Tables 1 and S3†).²⁴ It is reported in the case of carbon-supported Ru catalysts that a suitable subnanometer-sized metal cluster shows higher activity than larger- and smaller-sized metal particles.²⁵ Table 2 also shows that, when the pressure was increased to 2.0 MPa, the rate of the activated 1/HY increased by about twofold ($37.1 \text{ mmol g}_{\text{Mo}}^{-1} \text{ h}^{-1}$, entry 4). This rate was significantly higher than those of previously-reported supported Mo catalysts at the same reaction pressure even when different weight hourly space velocity (WHSV) values are considered: a silica-supported single-metal catalyst prepared from $\text{Mo}(\equiv\text{CBu}^t)(\text{Np})_3$ ($6.8 \text{ mmol g}_{\text{Mo}}^{-1} \text{ h}^{-1}$, entry 5)²⁶ and HZSM5-supported MoN_x ($4.3 \text{ mmol g}_{\text{Mo}}^{-1} \text{ h}^{-1}$, entry 6) as well

as MoC_x ($4.5 \text{ mmol g}_{\text{Mo}}^{-1} \text{ h}^{-1}$, entry 7) prepared by nitridation and carbonization of MoO_3 , respectively.²¹ The lower rate of supported MoN_x can be attributed to the weaker N_2 dissociation ability of the nitrided Mo surface than that of the metallic Mo surface.²⁷ Furthermore, our supported Mo metal clusters prepared from 1 have advantages in terms of the stability of the precursor in air and capability of ammonia synthesis after simple H_2 -activation using the same reaction tube, in contrast to the use of highly air-sensitive $\text{Mo}(\equiv\text{CBu}^t)(\text{Np})_3$ as a precursor or the need to transfer the sample from a quartz tube for air-calcination and succeeding nitridation or carbonization to a metal tube for ammonia synthesis under pressurized conditions.

The turnover frequency (TOF) of activated 1/HY (0.006, Table S4†) was 30% of that of a commercially-used Ru catalyst, Ba–Ru/C (0.02), based on the number of active metal sites.²⁸ As Table S5† shows, whereas the rate of the activated 1/HY per metal weight ($20.5 \text{ mmol g}_{\text{Mo}}^{-1} \text{ h}^{-1}$, entry 3) was higher than or comparable to that of an Fe catalyst ($<16 \text{ mmol g}_{\text{Fe}}^{-1} \text{ h}^{-1}$, entry 8) and about a quarter of that of an Ru catalyst ($91 \text{ mmol g}_{\text{Ru}}^{-1} \text{ h}^{-1}$, entry 9). The rate per catalyst weight ($0.483 \text{ mmol g}_{\text{cat}}^{-1} \text{ h}^{-1}$, entry 3) was much lower than those of the Fe ($14 \text{ mmol g}_{\text{cat}}^{-1} \text{ h}^{-1}$, entry 8) and Ru catalysts ($8.2 \text{ mmol g}_{\text{cat}}^{-1} \text{ h}^{-1}$, entry 9), under similar reaction conditions. The rate of 1/HY per catalyst weight allows a lot of room for increase by improvements such as the addition of promoters for industrialization.

The catalytic behavior of activated 1/HY with the highest synthesis rate among the supported Mo catalysts (Table 2, entry 3) was further investigated. As shown in Fig. S14,† the synthesis rate ($20.5 \text{ mmol g}_{\text{Mo}}^{-1} \text{ h}^{-1}$) remained constant for 258 h at 1.0 MPa, indicating that the activated supported cluster is highly durable for long-term ammonia synthesis. As shown in the Arrhenius plots (Fig. S15†), the apparent activation energy of activated 1/HY (89 kJ mol^{-1}) was lower than those of activated 1/HZSM5 (92 kJ mol^{-1}) and 1/MCM41 (110 kJ mol^{-1}), indicating that activated 1/HY is most effective for ammonia synthesis at lower temperatures. At a higher reaction pressure (5.0 MPa), 1/HY afforded ammonia catalytically even at 200 °C with a turnover number of more than 4 per Mo-atom (Fig. 3). Various catalysts with ammonia synthesis activity at 200 °C have been reported, with TOF values at around 1.0 MPa ranging in the

Table 2 Catalytic activities of various supported Mo catalysts for NH_3 synthesis at 400 °C

Entry	Catalyst (precursor)	Mo ratio (wt%)	NH_3 yield ^a ($\text{mmol g}_{\text{Mo}}^{-1} \text{ h}^{-1}$)	Reaction pressure	WHSV ($\text{mL g}_{\text{cat}}^{-1} \text{ h}^{-1}$)	Ref.
1	$(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (1)/MCM41	2.36	10.2 (0.7)	1.0 ^b MPa	18 000	This work
2	$(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (1)/HZSM5	2.36	14.7 (1.1)	1.0 ^b MPa	18 000	This work
3	$(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (1)/HY	2.36	20.5 (0.7)	1.0 ^b MPa	18 000	This work
4	$(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (1)/HY	2.36	37.1 (1.2)	2.0 ^b MPa	18 000	This work
5	$\text{Mo}(\equiv\text{CBu}^t)(\text{Np})_3/\text{SiO}_2$	2.0	6.8	2.0 MPa	12 000	26
6	$\text{MoN}_x/\text{HZSM5}^c$	2.17	4.3	2.0 MPa	9000	21
7	$\text{MoC}_x/\text{HZSM5}^c$	2.17	4.5	2.0 MPa	9000	21

^a The experiments were performed at least three times, and the values in parentheses are standard deviations. ^b Absolute pressure. ^c $\text{SiO}_2/\text{Al}_2\text{O}_3$ molar ratio = 70 (Si/Al molar ratio = 35).

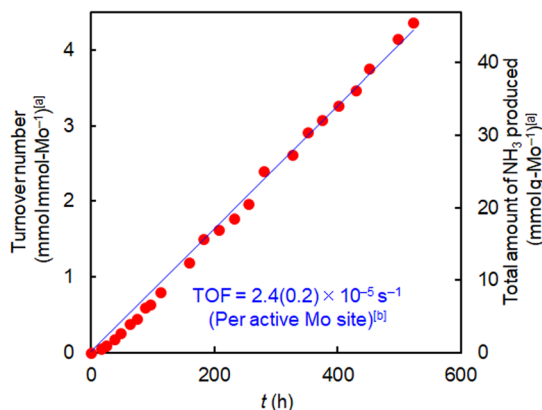


Fig. 3 Ammonia synthesis using $(\text{H}_3\text{O})_2[(\text{Mo}_6\text{Cl}_8)\text{Cl}_6] \cdot 6\text{H}_2\text{O}$ (**1**)/HY at 200 °C and 5.0 MPa (absolute pressure). Catalyst amount, 0.2 g; N_2/H_2 (1/3-mixture) flow rate, 60 mL min^{-1} . [a] Per Mo included in the catalysts. [b] Per NH_3 -adsorbing Mo (see Table S4†).

orders of 10^{-3} to 10^{-4} s^{-1} .^{29,30} In comparison with these values, the TOF of **1**/HY at 5.0 MPa ($2.4 \times 10^{-5} \text{ s}^{-1}$) (Fig. 3) was lower by one or two orders of magnitude. However, the ammonia synthesis rate of **1**/HY was stable for at least as long as 520 h, while the stability of the rate at 200 °C (for up to 26–480 h) has been reported for only a few catalysts.^{7d,31} In the case of **1**/HY, no pretreatment with a mixture of N_2 and H_2 at a higher temperature was necessary before the stable formation of ammonia at 200 °C, which is indicative of the high and sustainable N_2 dissociation ability of the supported Mo metal cluster at low temperatures.

Kinetic studies

To investigate the reaction mechanism for ammonia synthesis over the supported Mo metal clusters, kinetic and density functional theory (DFT) studies were carried out. In the kinetic studies, reaction orders with respect to N_2 , H_2 , and NH_3 for the clusters were measured (Fig. 4). The Mo metal cluster on HY as well as HZSM5 and MCM41 showed N_2 and H_2 reaction orders of around 0.5–0.6 and 1.0, respectively. This suggests that the reaction of dissociated N with H_2 forming N–H bonds is the rate-determining step (RDS), while the N_2 dissociation step is no

longer the RDS.³² In contrast, the RDS of the supported Mo catalysts reported previously (listed in entries 5 and 6 in Table 2) is ascribed to N_2 dissociation,^{21,26} as for most conventional catalysts.³³ Low N_2 reaction orders of around 0.5 have been reported for catalysts with electrides,^{28,31a,34} oxyhydrides,³⁵ a nitride-hydride,^{31c} and alkali- or alkaline earth-hydrides,^{29a,36} and the favorable N_2 dissociation has been attributed to strong electron donation by these catalyst components to the $\text{N}\equiv\text{N}$ bond. Our supported Mo metal clusters achieved low N_2 reaction orders without using such strongly electron-donating components.

The RDS of ammonia synthesis over the HY-supported cluster was investigated by comparing the experimental reaction rates with calculated ones. When the synthesis reaction is assumed to obey the Langmuir–Hinshelwood (dissociative) mechanism, in which cleavage of the N–N bond takes place before the formation of N–H bonds,^{33a} the eight elementary reaction steps are established (eqn (S6)–(S13)†). Among these eight steps, the dissociation of N_2 or the formation of NH , NH_2 , or NH_3 (eqn (S9)–(S12)†) can be attributed to the RDS because of its high activation energy, and the four corresponding calculated reaction rates can be deduced (eqn (S14)–(S17)†).^{29c,34b} These equations were fitted to a set of experimental rates obtained under various reaction gas ratios, using a least-squares method. Fig. S17† shows the best-fit of the calculated rates to the experimental ones. When the dissociation of N_2 was assumed to be the RDS, the fitting was poor with a negative value of determination coefficient. In contrast, when the formation of NH , NH_2 , and NH_3 was assumed to be the RDS, larger determination coefficient values (0.71–0.99) were obtained. These observations suggest that the RDS of ammonia synthesis over the HY-supported cluster is the formation step of NH , NH_2 , or NH_3 rather than the dissociation step of N_2 . This conclusion is also supported by the DFT results shown below.

DFT calculations

The reaction mechanism for the HY-supported cluster was studied in more detail using DFT calculations. As suggested by the XAFS and STEM analysis, after H_2 -activation and ammonia synthesis, the molybdenum cluster sizes were almost the same as that of hexanuclear precursor **1**. Thus, we assumed that the

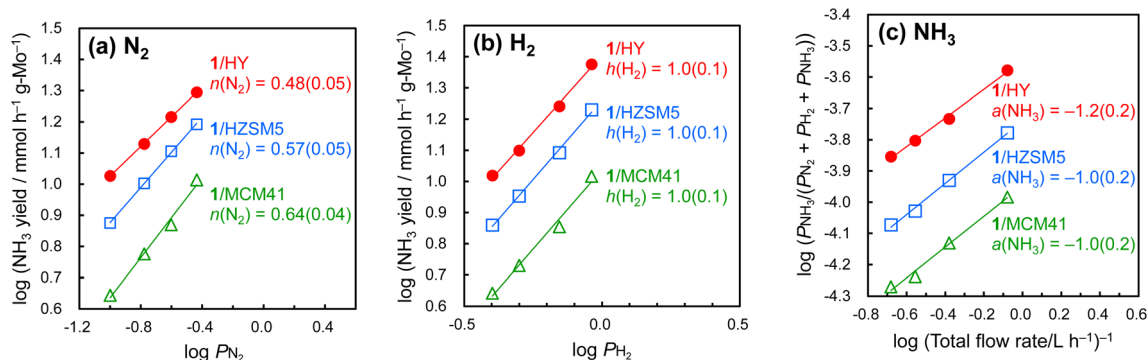


Fig. 4 Dependence of ammonia synthesis rate on the partial pressures of (a) N_2 , (b) H_2 , and (c) NH_3 over H_2 -activated clusters on various supports at 400 °C and 1 MPa.

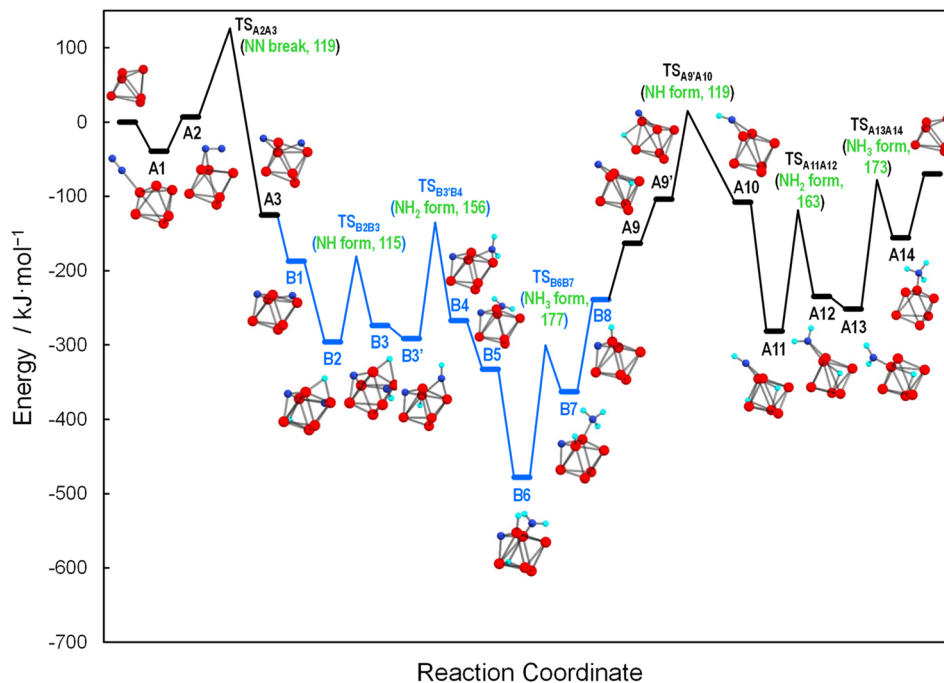


Fig. 5 Potential energy profiles and structural change along the dominant reaction pathway determined by microkinetic analysis. The HY zeolite and gaseous N_2 , H_2 , and NH_3 are omitted for clarity.

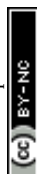
cluster during the ammonia synthesis also has a hexanuclear structure. We examined various local structures of HY zeolite to accommodate the Mo_6 metal cluster. When the cluster was located on the two adjacent four-membered oxygen rings with three Mo–oxygen interactions, the most stable structure was obtained (see Fig. S18 and Table S6† for details). Then, this model was used for further investigation of the reaction intermediates (Fig. S19†) and mechanism.

We determined several reaction pathways for ammonia synthesis through dissociative and associative mechanisms³⁷ (Fig. S20† and description therein), followed by microkinetic analysis using the potential energies of intermediates and transition states of these pathways (see Fig. S21† and description therein). Fig. 5 shows the dominant reaction pathway determined by microkinetic analysis. According to this, an N_2 molecule is first adsorbed on an Mo atom in a terminal end-on mode (A1), followed by a configurational change of the Mo-bonded N into the μ_2 -bridging adsorption mode (A2). Then, the N–N bond cleavage by participation of three Mo atoms takes place to afford two μ_2 -bridging N atoms (A3), followed by a configurational change of one N atom from μ_2 - to μ_3 -bridging (B1). Here, we use A and B notations to distinguish the pathway depending on the coordination mode of the N atom (μ_2 - or μ_3 -bridging) that is not involved in the first hydrogenation step (see Fig. S20† for details). The first H_2 molecule is subsequently introduced by dissociative adsorption (B2), followed by the transfer of one H to the μ_2 -bridging N to form a μ_2 -bridging NH (B3). After the migration of the remaining H atom (B3 $\rightarrow$ B3'), the second H transfers to NH to yield NH_2 (B4). After a minor configurational change of NH_2 from μ_2 -bridging to the terminal site (B4 $\rightarrow$ B5), the second H_2 molecule is introduced by

dissociative adsorption (B6). Then, the third H transfer to the NH_2 group affords a terminal NH_3 (B7), followed by release of the first NH_3 molecule (B8). After that, configurational change of the μ_3 -bridging N to a μ_2 -bridging mode takes place (A9). Further, after the migration of the H atom (A9 $\rightarrow$ A9'), the N atom accepts the fourth H atom to form a μ_2 -bridging NH (A10). After the third introduction of the H_2 molecule (A11), the fifth H transfer gives a μ_2 -bridging NH_2 (A12). Then, after the configurational change of the μ_2 -bridging NH_2 to terminal NH_2 (A12 $\rightarrow$ A13), the sixth H transfer occurs to afford a terminal NH_3 (A14). Finally, the second NH_3 molecule is released. In comparison with the energy barrier of N_2 dissociation (119 kJ mol^{-1} for TS_{A2A3}), the barriers of the first and second NH formation (115 kJ mol^{-1} for TS_{B2B3} and 119 kJ mol^{-1} for $TS_{A9'A10}$, respectively) are comparable, and those for the first and second NH_2 formation (156 kJ mol^{-1} for $TS_{B3'B4}$ and 163 kJ mol^{-1} for TS_{A11A12} , respectively) and the first and second NH_3 formation (177 kJ mol^{-1} for TS_{B6B7} and 173 kJ mol^{-1} for TS_{A13A14} , respectively) are higher. Thus, the RDS is not N–N bond cleavage but N–H bond formation, which is consistent with the experimental results. These results demonstrate that the angstrom-size Mo cluster prefers the dissociative pathway, in which the N–N bond cleavage promoted by the Mo_6 multinuclear structure is not the RDS.

Conclusions

In summary, supported angstrom-scale Mo metal clusters were prepared by the impregnation of a hexanuclear molecular halide cluster on various porous supports and subsequent activation with H_2 and were characterized by XAFS and STEM



analysis. When HY zeolite was used as the support, the Mo cluster size of precursor **1** was retained even after activation. The resulting angstrom-size metal cluster catalyzed ammonia synthesis from N₂ and H₂. The catalytic activity was highly durable even at 200 °C. The N≡N bond was effectively cleaved by the cooperation of multiple Mo sites, and the RDS shifted from N₂ dissociation to N–H formation, as confirmed by kinetic and computational studies. This work has expanded the scope of the application of a halide cluster for catalysis and developed a novel ultra-small Mo metal cluster catalyst for efficient ammonia synthesis, based on the multinuclearity of a metal cluster of suitable size.

Data availability

The data that support the findings of this study are available in the ESI† of this article.

Author contributions

Satoshi Kamiguchi: conceptualization (lead), data curation (lead), formal analysis (lead), funding acquisition (lead), investigation (lead), methodology (lead), project administration (lead), supervision (lead), validation (lead), and writing – original draft (lead). Kiyotaka Asakura: formal analysis (supporting), investigation (supporting), resources (lead), validation (supporting), and writing – original draft (supporting) on XAFS analysis. Tamaki Shibayama: formal analysis (supporting), investigation (supporting), resources (lead), validation (supporting), and writing – original draft (supporting) on STEM measurements. Tomoko Yokaichiya: formal analysis (supporting), investigation (lead), and writing – original draft (supporting) on DFT calculations. Tatsushi Ikeda: formal analysis (lead), investigation (supporting), and writing – original draft (supporting) on DFT calculations. Akira Nakayama: data curation (lead), funding acquisition (lead), project administration (lead), supervision (lead), validation (lead), and writing – original draft (lead) on DFT calculations. Ken-ichi Shimizu: formal analysis (supporting), resources (lead), investigation (supporting), validation (supporting), and writing – original draft (supporting) on catalyst analysis. Zhaomin Hou: conceptualization (supporting), project administration (supporting), supervision (supporting), validation (supporting), and writing – original draft (supporting).

Conflicts of interest

The authors declare no competing financial interest.

Acknowledgements

This work was supported by KAKENHI grant 19K05515 from the JSPS. This work was also supported by the Joint Usage/Research Center for Catalysis, Hokkaido University (Proposal No. 16B1016, 18B1012, 19B1023, and 20B1038). The calculations were partly performed on supercomputers at RCCS (Okazaki), RIIT (Kyushu Univ.), ACCMS (Kyoto Univ.), and the Center for Computational Materials Science, Institute for Materials

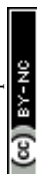
Research (Tohoku University, Proposal No. 202012-SCKXX-0002). The XAFS experiments were conducted at the BL551 and BL11S2 of Aichi Synchrotron Radiation Center, Aichi Science & Technology Foundation, Aichi, Japan (Proposal No. 201602001, 201905002, and 202002012). The STEM measurements were supported by the “Nanotechnology Platform Program” at Hokkaido University (A-19-HK-0055 and A-20-HK-0013). The authors thank the technical division of the Institute for Catalysis (Hokkaido University) for manufacturing glass cells for the XAFS measurements. The authors also thank the Materials Characterization Support Unit, RIKEN CEMS for the elemental analysis, and the MC Evolve Technologies Corporation for measurements of pore volume and BET surface area.

Notes and references

- (a) J. W. Erisman, M. A. Sutton, J. Galloway, Z. Klimont and W. Winiwarter, How a century of ammonia synthesis changed the world, *Nat. Geosci.*, 2008, **1**, 636–639; (b) J. M. Thomas and W. J. Thomas, *Principles and Practice of Heterogeneous Catalysis*, Wiley, Hoboken, 2014.
- (a) A. Klerke, C. H. Christensen, J. K. Nørskov and T. Vegge, Ammonia for hydrogen storage: challenges and opportunities, *J. Mater. Chem.*, 2008, **18**, 2304–2310; (b) F. Schüth, R. Palkovits, R. Schlögl and D. S. Su, Ammonia as a possible element in an energy infrastructure: catalysts for ammonia decomposition, *Energy Environ. Sci.*, 2012, **5**, 6278–6289; (c) J. Guo and P. Chen, Interplay of Alkali, Transition Metals, Nitrogen, and Hydrogen in Ammonia Synthesis and Decomposition Reactions, *Acc. Chem. Res.*, 2021, **54**, 2434–2444.
- D. Erdemir and I. Dincer, A perspective on the use of ammonia as a clean fuel: challenges and solutions, *Int. J. Energy Res.*, 2021, **45**, 4827–4834.
- L. E. Apodaca, *Mineral Commodity Summaries*, U.S. Department of the Interior, U.S. Geological Survey, 2022, pp. 126–127, available at <https://pubs.usgs.gov/periodicals/mcs2022/mcs2022.pdf>.
- (a) A. Mittasch, Early Studies of Multicomponent Catalysts, *Adv. Catal.*, 1950, **2**, 81–104; (b) C. W. Hooper, Ammonia synthesis: commercial practice, in *Catalytic Ammonia Synthesis, Fundamentals and Practice*, ed. J. R. Jennings, Springer, US, 1991, pp. 253–283.
- C. W. Hooper, Ammonia synthesis: commercial practice, in *Catalytic Ammonia Synthesis, Fundamentals and Practice*, ed. J. R. Jennings, Springer, US, 1991, pp. 303–348.
- (a) Y. V. Larichev, D. A. Shlyapin, P. G. Tsyrl'nikov and V. I. Bukhtiyarov, Comparative Study of Rubidium and Cesium as Promoters in Carbon-supported Ruthenium Catalysts for Ammonia Synthesis, *Catal. Lett.*, 2008, **120**, 204–209; (b) C. Fernández, C. Sassoye, D. P. Debecker, C. Sanchez and P. Ruiz, Effect of the size and distribution of supported Ru nano particles on their activity in ammonia synthesis under mild reaction conditions, *Appl. Catal., A*, 2014, **474**, 194–202; (c) C. Fernández, C. Pezzotta, E. M. Gaigneaux, N. Bion, D. Duprez and P. Ruiz, Disclosing the synergistic mechanism in the catalytic



- activity of different-sized Ru nanoparticles for ammonia synthesis at mild reaction conditions, *Catal. Today*, 2015, **251**, 88–95; (d) M. Hattori, S. Iijima, T. Nakao, H. Hosono and M. Hara, Solid solution for catalytic ammonia synthesis from nitrogen and hydrogen gases at 50 °C, *Nat. Commun.*, 2020, **11**, 2001; (e) K. H. R. Rouwenhorst, A. G. J. Van der Ham and L. Lefferts, Beyond Haber-Bosch: the renaissance of the Claude process, *Int. J. Hydrogen Energy*, 2021, **46**, 21566–21579.
- 8 C. C. Ahn and O. L. Krivanek, *EELS Atlas: A Reference Collection of Electron Energy Loss Spectra Covering All Stable Elements*, Gatan, Inc., 1983.
 - 9 (a) K. Aika, Role of alkali promoter in ammonia synthesis over ruthenium catalysts—Effect on reaction mechanism, *Catal. Today*, 2017, **286**, 14–20; (b) J. Humphreys, R. Lan and S. Tao, Development and Recent Progress on Ammonia Synthesis Catalysts for Haber-Bosch Process, *Adv. Energy Sustainability Res.*, 2021, **2**, 2000043; (c) H. Hosono and M. Kitano, Advances in Materials and Applications of Inorganic Electrides, *Chem. Rev.*, 2021, **121**, 3121–3185.
 - 10 (a) T. Shima, S. Hu, G. Luo, X. Kang, Y. Luo and Z. Hou, Dinitrogen Cleavage and Hydrogenation by a Trinuclear Titanium Polyhydride Complex, *Science*, 2013, **340**, 1549–1552; (b) T. Shima, G. Luo, S. Hu and Z. Hou, Experimental and Computational Studies of Dinitrogen Activation and Hydrogenation at a Tetranuclear Titanium Imide/Hydride Framework, *J. Am. Chem. Soc.*, 2019, **141**, 2713–2720; (c) T. Shima, J. Yang, G. Luo, Y. Luo and Z. Hou, Dinitrogen Activation and Hydrogenation by C₅Me₄SiMe₃-Ligated Di- and Trinuclear Chromium Hydride Complexes, *J. Am. Chem. Soc.*, 2020, **142**, 9007–9016.
 - 11 J. B. Howard and D. C. Rees, Structural Basis of Biological Nitrogen Fixation, *Chem. Rev.*, 1996, **96**, 2965–2982.
 - 12 (a) I. Dance, Elucidating the Coordination Chemistry and Mechanism of Biological Nitrogen Fixation, *Chem.-Asian J.*, 2007, **2**, 936–946; (b) J. Kästner and P. E. Blöchl, Ammonia Production at the FeMo Cofactor of Nitrogenase: Results from Density Functional Theory, *J. Am. Chem. Soc.*, 2007, **129**, 2998–3006.
 - 13 (a) S. Kamiguchi, S. Nagashima and T. Chihara, Characterization of Catalytically Active Octahedral Metal Halide Cluster Complexes, *Metals*, 2014, **4**, 84–107; (b) S. Nagashima, S. Kamiguchi and T. Chihara, Catalytic Reactions over Halide Cluster Complexes of Group 5–7 Metals, *Metals*, 2014, **4**, 235–313.
 - 14 S. Kamiguchi, T. Mori, M. Watanabe, A. Suzuki, M. Kodomari, M. Nomura, Y. Iwasawa and T. Chihara, Retention of the octahedral metal framework of Nb and Mo halide clusters in catalytic decomposition of phenyl acetate to phenol and ketene, *J. Mol. Catal. A*, 2006, **253**, 176–186.
 - 15 S. Nagashima, S. Kamiguchi, S. Ohguchi and T. Chihara, Gas-Phase Alkylation of Pyridine and Phenol with Alcohols Over Halide Clusters of Group 5–7 Transition Metals as Solid Acid Catalysts, *J. Cluster Sci.*, 2011, **22**, 647–660.
 - 16 (a) A. Logadottir, T. H. Rod, J. K. Nørskov, B. Hammer, S. Dahl and C. J. H. Jacobsen, The Brønsted–Evans–Polanyi Relation and the Volcano Plot for Ammonia Synthesis over Transition Metal Catalysts, *J. Catal.*, 2001, **197**, 229–231; (b) C. J. H. Jacobsen, S. Dahl, B. S. Clausen, S. Bahn, A. Logadottir and J. K. Nørskov, Catalyst Design by Interpolation in the Periodic Table: Bimetallic Ammonia Synthesis Catalysts, *J. Am. Chem. Soc.*, 2001, **123**, 8404–8405.
 - 17 (a) D. V. Yandulov and R. R. Schrock, Catalytic Reduction of Dinitrogen to Ammonia at a Single Molybdenum Center, *Science*, 2003, **301**, 76–78; (b) K. Arashiba, Y. Miyake and Y. Nishibayashi, *Nat. Chem.*, 2011, **3**, 120–125; (c) Y. Ashida, K. Arashiba, K. Nakajima and Y. Nishibayashi, A molybdenum complex bearing PNP-type pincer ligands leads to the catalytic reduction of dinitrogen into ammonia, *Nature*, 2019, **568**, 536–540.
 - 18 S. J. Pennycook, Z-contrast stem for materials science, *Ultramicroscopy*, 1989, **30**, 58–69.
 - 19 D. Hou, L. Grajciar, P. Nachtigall and C. J. Heard, Origin of the Unusual Stability of Zeolite-Encapsulated Sub-Nanometer Platinum, *ACS Catal.*, 2020, **10**, 11057–11068.
 - 20 W. Ding, S. Li, G. D. Meitzner and E. Iglesia, Methane Conversion to Aromatics on Mo/H-ZSM5: Structure of Molybdenum Species in Working Catalysts, *J. Phys. Chem. B*, 2001, **105**, 506–513.
 - 21 N. Liu, L. Nie, N. Xue, H. Dong, L. Peng, X. Guo and W. Ding, Catalytic Ammonia Synthesis over Mo Nitride/ZSM-5, *ChemCatChem*, 2010, **2**, 167–174.
 - 22 Y. Iwasawa and M. Yamaguchi, New SiO₂-Attached Mo-Pair Catalysts—preparations, Surface Structures, and Chemical Nature, *J. Catal.*, 1983, **82**, 373–381.
 - 23 C. Aydin, J. Lu, M. Shirai, N. D. Browning and B. C. Gates, Ir₆ Clusters Compartmentalized in the Supercages of Zeolite NaY: Direct Imaging of a Catalyst with Aberration-Corrected Scanning Transmission Electron Microscopy, *ACS Catal.*, 2011, **1**, 1613–1620.
 - 24 (a) A. Beniya, S. Higashi, N. Ohba, R. Jinnouchi, H. Hirata and Y. Watanabe, CO oxidation activity of non-reducible oxide-supported mass-selected few-atom Pt single-clusters, *Nat. Commun.*, 2020, **11**, 1888; (b) A. Corma, P. Concepción, M. Boronat, M. J. Sabater, J. Navas, M. J. Yacaman, E. Larios, A. Posadas, M. A. López-Quintela, D. Buceta, E. Mendoza, G. Guilera and A. Mayoral, Exceptional oxidation activity with size-controlled supported gold clusters of low atomicity, *Nat. Chem.*, 2013, **11**, 775.
 - 25 (a) Y. Zhou, Q. Sai, Z. Tan, C. Wang, X. Wang, B. Lin, J. Ni, J. Lin and L. Jiang, Highly efficient subnanometer Ru-based catalyst for ammonia synthesis via an associative mechanism, *Chin. J. Chem. Eng.*, 2022, **43**, 177–184; (b) L. Lin, Y.-F. Jiang, T. Zhang, H. Cai, Y. Zhou, B. Lin, X. Lin, Y. Zheng, L. Zheng, X. Wang, C.-Q. Xu, C. Au, L. Jiang and J. Li, Size sensitivity of supported Ru catalysts for ammonia synthesis: from nanoparticles to subnanometric clusters and atomic clusters, *Chem*, 2022, **8**, 749–768.
 - 26 L. M. Azofra, N. Morlanes, A. Poater, M. K. Samantaray, B. Vidjayacoumar, K. Albahily, L. Cavallo and J.-M. Basset, Single-Site Molybdenum on Solid Support Materials for



- Catalytic Hydrogenation of N_2 -into- NH_3 , *Angew. Chem., Int. Ed.*, 2018, **57**, 15812–15816.
- 27 J. Zhao, C. Cui, H. Wang, J. Han, X. Zhu and Q. Ge, Anisotropic N-Modification of the Mo_4 Ensemble for Efficient Ammonia Synthesis on Molybdenum Nitrides, *J. Phys. Chem. C*, 2020, **124**, 616–624.
- 28 M. Kitano, Y. Inoue, Y. Yamazaki, F. Hayashi, S. Kanbara, S. Matsuishi, T. Yokoyama, S.-W. Kim, M. Hara and H. Hosono, Ammonia synthesis using a stable electride as an electron donor and reversible hydrogen store, *Nat. Chem.*, 2012, **4**, 934–940.
- 29 (a) W. Gao, P. Wang, J. Guo, F. Chang, T. He, Q. Wang, G. Wu and P. Chen, Barium Hydride-Mediated Nitrogen Transfer and Hydrogenation for Ammonia Synthesis: A Case Study of Cobalt, *ACS Catal.*, 2017, **7**, 3654–3661; (b) P. Wang, H. Xie, J. Guo, Z. Zhao, X. Kong, W. Gao, F. Chang, T. He, G. Wu, M. Chen, L. Jiang and P. Chen, The Formation of Surface Lithium–Iron Ternary Hydride and its Function on Catalytic Ammonia Synthesis at Low Temperatures, *Angew. Chem., Int. Ed.*, 2017, **56**, 8716–8720; (c) M. Kitano, Y. Inoue, M. Sasase, K. Kishida, Y. Kobayashi, K. Nishiyama, T. Tada, S. Kawamura, T. Yokoyama, M. Hara and H. Hosono, Self-organized Ruthenium–Barium Core–Shell Nanoparticles on a Mesoporous Calcium Amide Matrix for Efficient Low-Temperature Ammonia Synthesis, *Angew. Chem., Int. Ed.*, 2018, **57**, 2648–2652; (d) K. Sato, S. Miyahara, K. Tsujimaru, Y. Wada, T. Toriyama, T. Yamamoto, S. Matsumura, K. Inazu, H. Mohri, T. Iwasa, T. Taketsugu and K. Nagaoka, Barium Oxide Encapsulating Cobalt Nanoparticles Supported on Magnesium Oxide: Active Non-Noble Metal Catalysts for Ammonia Synthesis under Mild Reaction Conditions, *ACS Catal.*, 2021, **11**, 13050–13061; (e) S. Miyahara, K. Sato, K. Tsujimaru, Y. Wada, Y. Ogura, T. Toriyama, T. Yamamoto, S. Matsumura, K. Inazu and K. Nagaoka, Co Nanoparticle Catalysts Encapsulated by $BaO-La_2O_3$ Nanofractions for Efficient Ammonia Synthesis Under Mild Reaction Conditions, *ACS Omega*, 2022, **7**, 24452–24460; (f) Z. Li, Y. Lu, J. Li, M. Xu, Y. Qi, S.-W. Park, M. Kitano, H. Hosono, J.-S. Chen and T.-N. Ye, Multiple reaction pathway on alkaline earth imide supported catalysts for efficient ammonia synthesis, *Nat. Commun.*, 2023, **14**, 6373; (g) K. Era, K. Sato, S. Miyahara, T. Naito, K. De Silva, S. Akrami, H. Yamada, T. Toriyama, T. Yamamoto, Y. Murakami, K. Aika, K. Inazu and K. Nagaoka, Catalytic Behavior of K-doped Fe/MgO Catalysts for Ammonia Synthesis Under Mild Reaction Conditions, *ChemSusChem*, 2023, **16**, e202300942.
- 30 When the TOF at a different temperature (TOF_T) was reported, the TOF at 200 °C (TOF_{200}) was calculated using the ammonia synthesis rates at these temperatures (r_T and r_{200}) as follows: $TOF_{200} = TOF_T \times (r_{200}/r_T)$.
- 31 (a) Y. Inoue, M. Kitano, K. Kishida, H. Abe, Y. Niwa, M. Sasase, Y. Fujita, H. Ishikawa, T. Yokoyama, M. Hara and H. Hosono, Efficient and Stable Ammonia Synthesis by Self-Organized Flat Ru Nanoparticles on Calcium Amide, *ACS Catal.*, 2016, **6**, 7577–7584; (b) W. Gao, S. Feng, H. Yan, Q. Wang, H. Xie, L. Jiang, W. Zhang, Y. Guan, H. Wu, H. Cao, J. Guo and P. Chen, In situ formed Co from a Co–Mg–O solid solution synergizing with LiH for efficient ammonia synthesis, *Chem. Commun.*, 2021, **57**, 8576–8579; (c) Y. Guan, W. Zhang, Q. Wang, C. Weidenthaler, A. Wu, W. Gao, Q. Pei, H. Yan, J. Cui, H. Wu, S. Feng, R. Wang, H. Cao, X. Ju, L. Liu, T. He, J. Guo and P. Chen, Barium chromium nitride-hydride for ammonia Synthesis, *Chem Catal.*, 2021, **1**, 1042–1054; (d) M. Hattori, N. Okuyama, H. Kurosawa and M. Hara, Low-Temperature Ammonia Synthesis on Iron Catalyst with an Electron Donor, *J. Am. Chem. Soc.*, 2023, **145**, 7888–7897.
- 32 M. Kitano, S. Kanbara, Y. Inoue, N. Kuganathan, P. V. Sushko, T. Yokoyama, M. Hara and H. Hosono, Electride support boosts nitrogen dissociation over ruthenium catalyst and shifts the bottleneck in ammonia synthesis, *Nat. Commun.*, 2015, **6**, 6731.
- 33 (a) K. Aika, M. Kumasaki, T. Oma, O. Kato, H. Matsuda, N. Watanabe, K. Yamazaki, A. Ozaki and T. Onishi, Support and Promoter Effect of Ruthenium Catalyst. III. Kinetics of Ammonia Synthesis over Various Ru Catalysts, *Appl. Catal., A*, 1986, **28**, 57–68; (b) F. Rosowski, A. Hornung, O. Hinrichsen, D. Herein, M. Muhler and G. Ertl, *Appl. Catal., A*, 1997, **151**, 443–460; (c) S. Hagen, R. Barfod, R. Fehrmann, C. J. H. Jacobsen, H. T. Teunissen and I. Chorkendorff, Ruthenium catalysts for ammonia synthesis at high pressures: preparation, characterization, and power-law kinetics, *J. Catal.*, 2003, **214**, 327–335; (d) S. E. Siporin and R. J. Davis, Use of kinetic models to explore the role of base promoters on Ru/MgO ammonia synthesis catalysts, *J. Catal.*, 2004, **225**, 359–368.
- 34 (a) M. Kitano, Y. Inoue, H. Ishikawa, K. Yamagata, T. Nakao, T. Tada, S. Matsuishi, T. Yokoyama, M. Hara and H. Hosono, Essential role of hydride ion in ruthenium-based ammonia synthesis catalysts, *Chem. Sci.*, 2016, **7**, 4036–4043; (b) Y. Gong, J. Wu, M. Kitano, J. Wang, T.-N. Ye, J. Li, Y. Kobayashi, K. Kishida, H. Abe, Y. Niwa, H. Yang, T. Tada and H. Hosono, Ternary intermetallic $LaCoSi$ as a catalyst for N_2 Activation, *Nat. Catal.*, 2018, **1**, 178–185; (c) H. Mizoguchi, M. Okunaka, M. Kitano, S. Matsuishi, T. Yokoyama and H. Hosono, Hydride-Based Electride Material, LnH_2 ($Ln = La, Ce, \text{ or } Y$), *Inorg. Chem.*, 2016, **55**, 8833–8838.
- 35 Y. Tang, Y. Kobayashi, N. Masuda, Y. Uchida, H. Okamoto, T. Kageyama, S. Hosokawa, F. Layer, K. Mitsuhara, K. Yamanaka, Y. Tamenori, C. Tassel, T. Yamamoto, T. Tanaka and H. Kageyama, Metal-Dependent Support Effects of Oxyhydride-Supported Ru, Fe, Co Catalysts for Ammonia Synthesis, *Adv. Energy Mater.*, 2018, **8**, 1801772.
- 36 P. K. Wang, F. Chang, W. B. Gao, J. P. Guo, G. T. Wu, T. He and P. Chen, Breaking scaling relations to achieve low-temperature ammonia synthesis through LiH-mediated nitrogen transfer and hydrogenation, *Nat. Chem.*, 2017, **9**, 64–70.
- 37 C. D. Zeinalipour-Yazdi, J. S. Hargreaves and C. R. A. Catlow, Low-T Mechanisms of Ammonia Synthesis on Co_3Mo_3N , *J. Phys. Chem. C*, 2018, **122**, 6078–6082.
- 38 W. C. Hamilton, Significance Tests on the Crystallographic R Factor, *Acta Crystallogr.*, 1965, **18**, 502–510.

