

RESEARCH ARTICLE

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11, 5244**Electronic metal–support interaction via Ni defect-induced Ru-modified Ni–CeO₂ for enhanced hydrogen oxidation activity†**Shuqing Zhou,^a Yi Liu,^a Lianrui Cheng,^a Tayirjan Taylor Isimjan,^{*b} Jianniao Tian ^{*a}
and Xiulin Yang ^{*a}

The creation of refined surface vacancies, a crucial bridge between theoretical structural studies and catalyst design, has attracted significant attention. Herein, we utilize a MOF pyrolytic derivatization strategy to create monodispersed Ru nanoparticles anchored on Ni–CeO₂ mesoporous microspheres with abundant Ni vacancies (Ru/Ni–CeO₂). Introducing nickel vacancies in Ru/Ni–CeO₂ is conducive to enhancing electrical conductivity and accelerating mass-charge transfer efficiency. As anticipated, the Ru/Ni–CeO₂ displays admirable hydrogen oxidation reaction (HOR) electrocatalytic activity with exchange current density (j_0) and mass activity reaching 3.27 mA cm⁻² and 1.93 mA g_{Ru}⁻¹, respectively, surpassing the values for cutting-edge Pt/C and most recorded Ru-based HOR electrocatalysts. Surprisingly, Ru/Ni–CeO₂ demonstrates robust tolerance to 1000 ppm CO, outperforming Pt/C. Integrated analysis suggests that Ni defect-induced directional electron transfer at the Ru/Ni–CeO₂ heterointerface arises from a strong electron–metal support interaction (EMSI) effect between Ru and Ni–CeO₂. This interaction optimizes the adsorption of H and OH, thereby enhancing HOR behavior.

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rsc.li/frontiers-inorganic**Introduction**

Hydrogen-based fuel cell technology, a promising avenue for hydrogen energy conversion, has experienced rapid development.¹ The HOR is one of the key reactions in hydrogen-oxygen fuel cells, acting as an intermediary between the electricity and hydrogen sectors and greatly affecting the conversion efficiency. Although proton exchange membrane fuel cells (PEMFCs) hit a pinnacle in their progression, alkaline exchange membrane fuel cells (AEMFCs) have increasingly supplanted PEMFCs because they can utilize non-precious metal materials to drive the cathode oxygen reduction reaction (ORR).² Unfortunately, the kinetics of the anodic HOR decreases dramatically in alkaline media, dropping at least two orders of magnitude relative to acidic conditions, severely hampering its commercialization.³ Based on Sabatier's principle, most active HOR electrocatalysts are positioned at the

summit of the volcano diagram, characterized by a moderate binding energy for adsorbed hydrogen (H_{ads}).⁴ Moreover, hydroxyl binding energy (OHBE) changes can elucidate the transformation of the HOR rate-determining step, and optimization OHBE can promote the Volmer step, which is essential for improving the HOR kinetics under alkaline electrolytes.⁵ Consequently, striving to harmonize the ideal bifunctional mechanism of hydrogen binding energy (HBE) and OHBE is an effective strategy to guide the optimization of alkaline HOR electrocatalysts.⁶

Ruthenium (Ru) is regarded as the most promising candidate for replacing platinum owing to its abundant reserves, reasonable Ru–H bond strength, and more affordable price than platinum.⁷ Incorporating ruthenium into nickel-based catalysts can tune the surface electronic structure of the catalyst, thereby optimizing the hydrogen binding energy (HBE) and achieving excellent HOR performance.⁸ For instance, Zhang *et al.* introduced Ru into Ni-based catalysts, and revealed that the RuNi alloy could effectively reduce the HBE, thus increasing the catalytic rate of the HOR.⁹ Furthermore, the adsorption of hydroxyl (OH_{ad}) also plays a crucial role in the HOR in alkaline media, as OH⁻ is involved in the fundamental steps.¹⁰ Sun *et al.* confirmed that the strength of OH adsorption at the Ru/RuO₂ interface was better than that of the comparison sample through the Ru/RuO₂ heterostructure, suggesting that the enhanced OH_{ads} binding energy was

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favorable for the HOR reaction.¹¹ Weakening HBE and enhancing OHBE are inferred as the primary reasons for improved HOR performance. Another viable method to enhance the HOR activity is to construct an electronic metal-support interaction (EMSI) within the catalyst, which can effectively modulate the electronic structure between the metal and the support, thereby significantly improving catalytic activity. Recently, Yang *et al.* reported a Ru/Ni-NiO@C electrocatalyst constructed using a MOF-engaged replacement-pyrolysis strategy, suggesting that the strong EMSI effect between Ru and Ni-NiO could effectively improve the HOR performance.¹² Consequently, the construction of Ru-based catalysts with attenuated HBE and enhanced OHBE is of great significance for guidance of the exploitation of high-efficiency HOR catalysts.

This study employed a MOF-assisted strategy and *in situ* modified Ru on three-dimensional mesoporous carbon nanospheres to fabricate a heterogeneous catalyst (Ru/Ni-CeO₂) with high HOR activity. The optimized Ru/Ni-CeO₂ catalysts exhibited an open structure and abundant nickel defects, showcasing high activity and exceptional CO tolerance, attrib-

ted to the enhanced EMSI effect between Ru and Ni-CeO₂. Systematic catalyst characterization (EPR, UPS, and XPS *etc.*) revealed that Ni defect-induced Ru-modified Ni-CeO₂ underwent obvious charge remodeling, leading to enhanced charge transport and accelerated adsorption and desorption of active intermediates. This work provides new inspiration for building efficient and CO-resistant HOR catalysts *via* defect-engineered induced EMSI effects.

Results and discussion

Synthesis and characterization

As depicted in Fig. 1a, we put forward a MOF-engaged pyrolysis-reduction approach for manufacturing Ru-modified Ni-CeO₂ microspheres. Specifically, Ce(NO₃)₃·6H₂O, Ni(NO₃)₂·6H₂O, and H₃BTC precursors were mixed in certain proportions by hydrothermal and high-temperature pyrolysis strategies to prepare Ni-CeO₂ microspheres with abundant nickel vacancies. Afterwards, the Ru precursor (RuCl₃) was confined

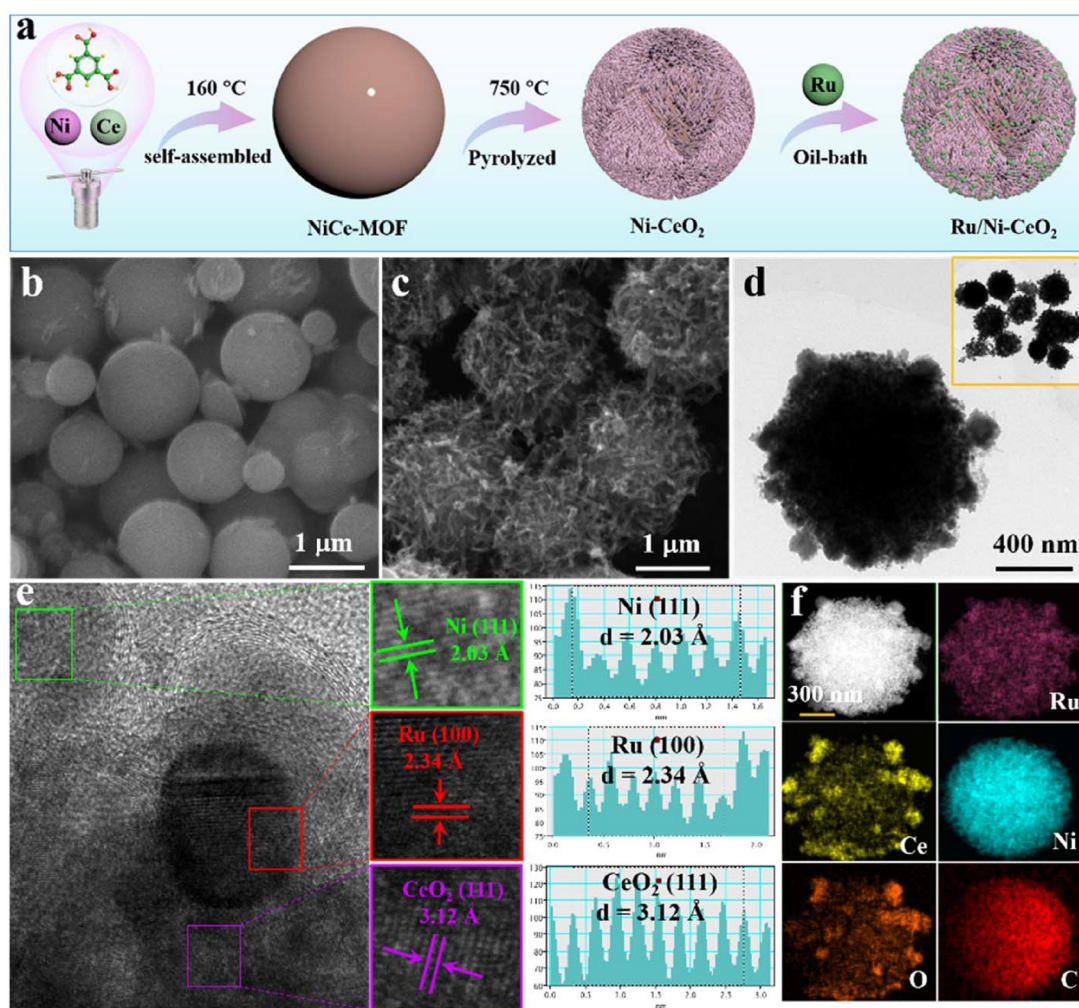


Fig. 1 (a) Preparation diagram of Ru/Ni-CeO₂ catalyst. SEM images of (b) NiCe-MOF and (c) Ru/Ni-CeO₂. (d) TEM image of Ru/Ni-CeO₂. (e) High-resolution TEM image of Ru/Ni-CeO₂. (f) Elemental mappings of Ru/Ni-CeO₂.

to the outer shells of the Ni–CeO₂ mesoporous spheres, forming a Ru/Ni–CeO₂ hybrid structure. The ample specific surface area of the finely spherical structure facilitated Ru dispersion, exposing more active sites.¹³ Correspondingly, Ru/Ni and Ru/CeO₂ were prepared employing a similar approach. Scanning electron microscopy (SEM) and transmission electron microscopy (TEM) were performed to explore the morphology of the catalyst. As shown in Fig. 1b, NiCe-MOF was dispersed spherically with a smooth surface. After high-temperature annealing, smooth spherical surfaces transformed into a nanonetwork structure (Fig. 1c). Furthermore, TEM observation verified that the morphology was well inherited after modifying Ru onto Ni–CeO₂ (Fig. 1d). High-resolution TEM (HRTEM) image examination (Fig. 1e) exhibited lattice fringes for the Ru/Ni–CeO₂ microspheres of 2.03, 2.34 and 3.12 Å, representing the (111) plane of Ni, the (100) plane of Ru, and the (111) plane of CeO₂, respectively.¹⁴ The coexistence of Ni, Ru, and CeO₂ components confirmed the successful fabrication of Ru/Ni–CeO₂ catalysts. Elemental mapping analysis presented in Fig. 1f verified the uniform distribution of Ru, Ce, Ni, O, and C elements throughout the Ru/Ni–CeO₂ microsphere structure.

XRD analysis was conducted to examine the crystal structure of the catalyst. As depicted in Fig. 2a, the diffraction peaks of Ru/Ni–CeO₂ matched with cubic phase CeO₂ (JCPDS: 78-0694) and cubic phase Ni (JCPDS: 87-0712), respectively. Similarly, Ru/CeO₂ and Ru/Ni exhibited matching peaks with cubic CeO₂ and cubic Ni (Fig. S1†), confirming the successful formation of the catalysts. However, no significant Ru diffraction peaks were observed in the XRD patterns, probably owing

to the amorphous state of Ru or the minimal Ru content.¹⁵ Furthermore, the electron paramagnetic resonance (EPR) test was applied to verify the existence of nickel vacancies. Apparently, almost symmetrical EPR signals around $g = 2.26$ were obtained in Ru/Ni–CeO₂ and Ru/Ni, while they disappeared in Ru/CeO₂ (Fig. 2b), which implied that there were Ni vacancies in Ru/Ni–CeO₂ and Ru/Ni that trapped unpaired electrons.¹⁶ Additionally, the nickel vacancies were investigated at different calcination temperatures and varying Ru content samples (Fig. S2†). Identical EPR signals were found in these samples, indicating the presence of Ni vacancies. The Ru/Ni–CeO₂ calcinated at 700 °C with a Ru content of 3.75 wt% exhibited a more pronounced EPR signal, suggesting a greater nickel vacancy concentration. The rich Ni vacancies in Ru/Ni–CeO₂ can enhance the concentration of charge carriers and the synergistic effect between the components, promoting the adsorption and desorption of intermediates.¹⁷

Raman spectroscopy was employed to assess the nano-materials' graphitization degree and structural defects.¹⁸ Typically, broad D- (~1345 cm⁻¹) and G- (~1597 cm⁻¹) bands reflect the disorder/edge defects and graphitization of the carbon species, respectively (Fig. 2c).¹³ The I_D/I_G values of Ru/Ni–CeO₂ (0.68) were the largest compared with those of Ru/Ni (0.62) and Ru/CeO₂ (0.57), suggesting that Ru/Ni–CeO₂ harvested more edge defects, favoring the intake of diverse active sites.¹⁹ Remarkably, the Raman spectra of Ru/Ni–CeO₂ and Ru/Ni exhibited prominent peaks around 500 cm⁻¹, but they were absent in the spectra of Ru/CeO₂. Based on previous investigations, the Raman modes at about 500 cm⁻¹ were inactive in perfect single crystals; however, surface effects or the

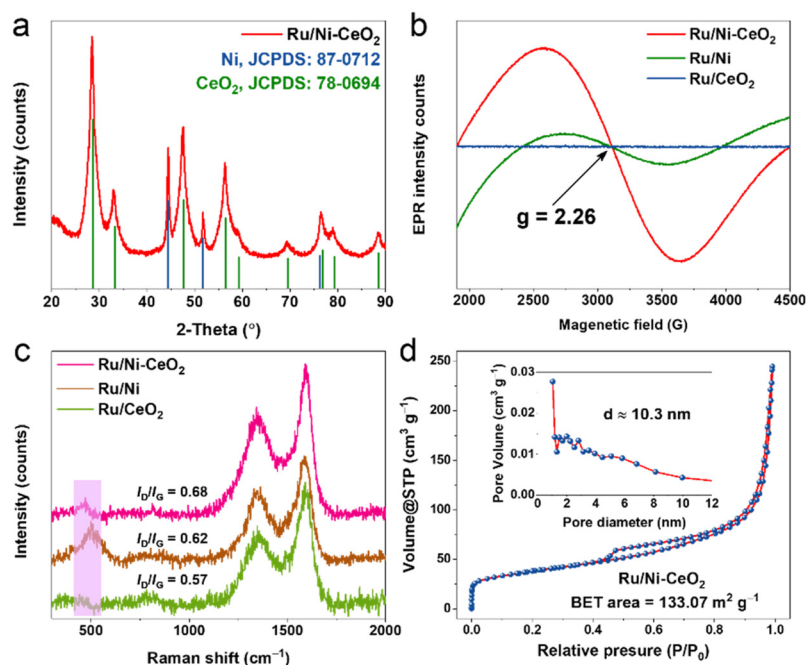


Fig. 2 (a) XRD pattern of Ru/Ni–CeO₂. (b) EPR spectra of Ru/Ni, Ru/CeO₂ and Ru/Ni–CeO₂. (c) Raman spectra of Ru/Ni, Ru/CeO₂ and Ru/Ni–CeO₂. (d) N₂ adsorption–desorption isotherms with the corresponding pore size distribution of Ru/Ni–CeO₂.

existence of vacancies made them apparent.¹⁸ This further confirmed the appearance of a large number of intrinsic nickel vacancies in Ru/Ni–CeO₂ and Ru/Ni.²⁰ The abundant nickel vacancies in Ru/Ni–CeO₂ dramatically tuned the local electronic structure, thus increasing the charge-carrier concentration, accelerating charge transfer and improving the catalytic activity.¹⁷ N₂ adsorption–desorption isotherms are displayed in Fig. 2d. Ru/Ni–CeO₂ exhibited an IV-type curve with a visible hysteresis loop, indicating that the Ru/Ni–CeO₂ nanospheres possessed a mesoporous structure with a pore size distribution of approximately 10.3 nm. Additionally, Ru/Ni–CeO₂ possessed the highest Brunauer–Emmett–Teller (BET) surface area of 133.07 m² g^{−1} compared to Ru/Ni (38.24 m² g^{−1}) and Ru/CeO₂ (54.77 m² g^{−1}), which contributed to the increased accessibility of the active site (Fig. S3†). The abundant specific surface area of the Ru/Ni–CeO₂ catalyst favored the dispersion of catalytic active sites, and the mesoporous structure was conducive to the adsorption of H₂ molecules and electron/mass transfer at heterogeneous interfaces, enhancing its catalytic performance.²¹

X-ray photoelectron spectroscopy (XPS) was employed to elucidate the surface chemical state and electronic structure of materials to identify the chemical environments of atoms. The XPS survey spectrum of Ru/Ni–CeO₂ verified the coexistence of

C, O, Ni, Ce, and Ru elements (Fig. S4†). High-resolution (HR) XPS spectra of C 1s + Ru 3d (Fig. S5†) displayed deconvolution peaks at 284.0, 284.8, 286.0, 288.2, and 290.1 eV, corresponding to C=C, C–C, C–O, C=O, and O–C=O, respectively, serving as calibration benchmarks.^{19,22} Meanwhile, deconvolution peaks near 281.0 and 285.1 eV represented 3d_{5/2} and 3d_{3/2} of metal Ru, respectively,²³ and the peaks at about 282.4 and 286.6 eV were 3d_{5/2} and 3d_{3/2} of RuO₂, respectively, owing to unavoidable oxidation by exposure to air.²⁴ Additionally, the high-resolution XPS spectra of Ru 3p further revealed the valence state of Ru. Fig. 3a depicts the deconvolution peaks at approximately 462.8 and 485.2 eV, corresponding to the 3p_{3/2} and 3p_{1/2} orbitals of metal Ru, respectively. The peaks situated around 465.9 and 488.2 eV are attributed to the 3p_{3/2} and 3p_{1/2} orbitals of RuO₂, which aligns with the literature reports.²⁵ Notably, the binding energy of Ru species in Ru/Ni–CeO₂ was lower than that in Ru/CeO₂, indicating electronic interaction between Ru and Ni–CeO₂.¹⁴ In Fig. 3b, four deconvoluted peaks at 853.3, 856.9, 859.5 and 863.2 eV were associated with Ni⁰, Ni²⁺, Ni³⁺ and satellite peak species, respectively.¹² The appearance of Ni²⁺ species probably resulted from the strong interaction of unreduced Ni²⁺ with CeO₂.²⁶ Meanwhile, the existence of Ni³⁺ signified that Ni vacancies were successfully integrated into the lattice of Ru/Ni–CeO₂ in the annealing

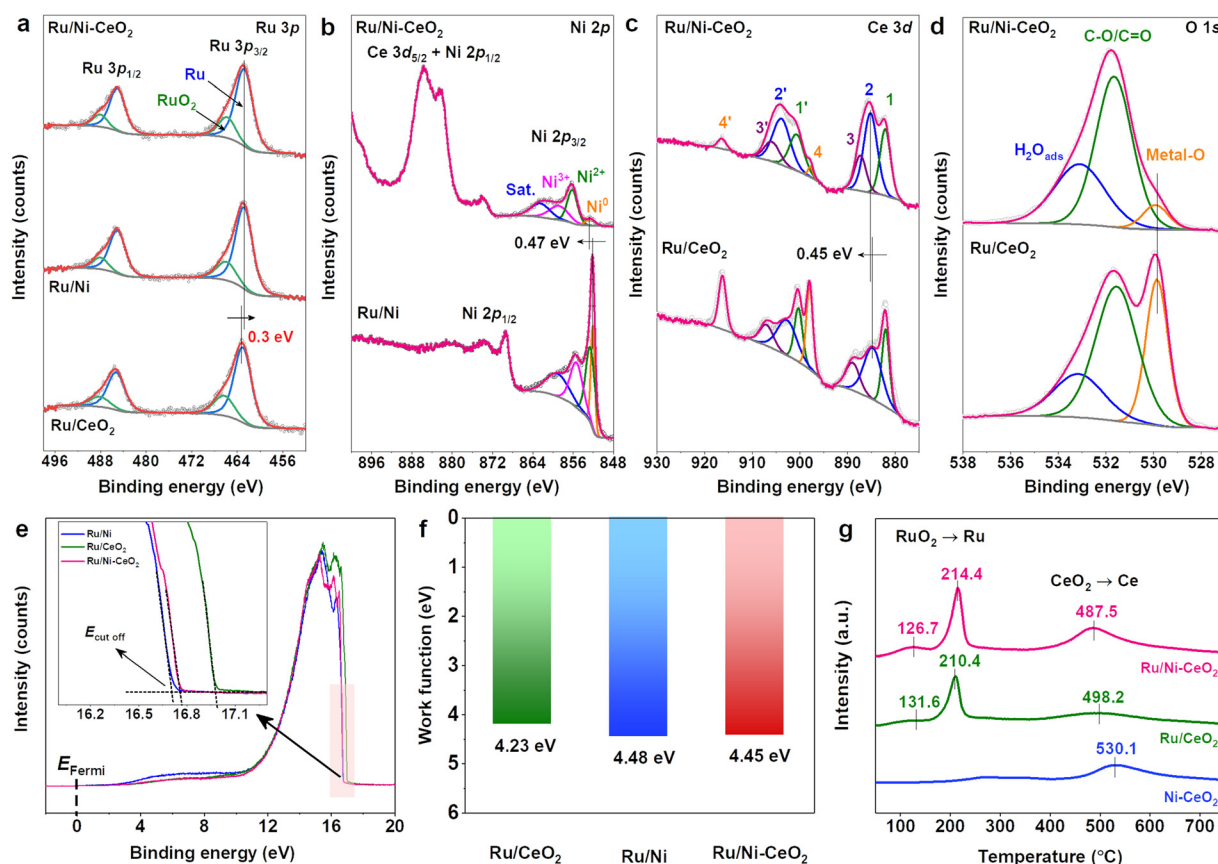


Fig. 3 High-resolution XPS spectra of (a) Ru 3p, (b) Ni 2p, (c) Ce 3d, and (d) O 1s of catalysts. (e) UPS spectra with inset partial enlargement of Ru/Ni, Ru/CeO₂ and Ru/Ni–CeO₂ catalysts and (f) corresponding work function. (g) H₂-TPR profiles of Ni–CeO₂, Ru/CeO₂ and Ru/Ni–CeO₂ catalysts.

process.¹² The oxidation of Ni²⁺ to Ni³⁺ helped to maintain the charge balance of the near-surface Ni vacancies, in line with the above Raman results.¹⁸ Intriguingly, the Ni⁰ peak in Ru/Ni–CeO₂ exhibited a positive shift of 0.47 eV compared to that in Ru/Ni, revealing a drastic electronic coupling between Ni and the other components.²⁶ Charge redistribution in the Ru/Ni–CeO₂ catalyst accelerated charge transfer between components, enhancing catalytic performance.²⁷

As displayed in Fig. 3c, the Ce 3d spectra of the catalysts could be deconvoluted into 8 peaks. The distinctive peaks observed at 885.9 and 904.6 eV (2, 2') were associated with Ce³⁺ (3d¹⁰4f¹ initial electronic state), whereas the six peaks detected at 882.8, 888.1, 898.5 (1, 3, 4), and 901.4, 906.6, 917.1 eV (1', 3', 4') were linked to Ce⁴⁺ (3d¹⁰4f⁰ initial electronic state).²⁸ Ce³⁺ and Ce⁴⁺ were identified in the Ru/Ni–CeO₂ catalyst, with Ce³⁺ likely resulting from electronic interaction between Ni²⁺ and CeO₂.²⁶ Apparently, the peak of Ce³⁺ in Ru/Ni–CeO₂ exhibited a positive shift of approximately 0.45 eV compared to Ru/CeO₂, indicating an electrostatic interaction between the CeO₂ and the surrounding Ni.²⁹ Deconvolution peaks of the O 1s XPS spectra (Fig. 3d) displayed three peaks at 529.8, 531.5, and 533.1 eV, matching metal–O, C–O/C=O, and adsorbed H₂O, respectively.³⁰ Additionally, the work function (WF) examined by ultraviolet photoelectron spectroscopy (UPS) can reveal the surface electronic properties of catalysts.³¹ As depicted in Fig. 3e and f, the WF values of Ru/CeO₂, Ru/Ni and Ru/Ni–CeO₂ were 4.23, 4.48 and 4.45 eV, respectively. Differences in WF among catalysts corresponded to variations in surface valence states, forming unique empty d-orbitals and accelerating charge inter-transfer, consistent with XPS results.³

Temperature-programmed reduction (H₂-TPR) is an effective approach for probing the interfacial interactions between metals and support. As displayed in Fig. 3g, shoulder peaks were observed at 126.7 and 131.6 °C, which were induced by oxygen adsorption.³² A peak at 214.4 °C was associated with RuO_x species strongly interacting with CeO₂.³³ The third peak at 487.5 °C was attributed to the surface oxygen reduction of CeO₂ away from Ru species. Compared to pure Ni–CeO₂ (530.1 °C), a dramatic decrease in the reduction temperature of Ru/Ni–CeO₂ (487.5 °C) was due to the loosening of Ce–O bonds strongly bound to Ru species.³² For Ru/CeO₂ (498.2 °C), the higher reduction temperature of CeO₂ implied a stronger EMSI effect of Ru/Ni–CeO₂ compared to Ru/CeO₂. The robust EMSI effect of Ru/Ni–CeO₂ facilitated the rapid adsorption and desorption of intermediates, accelerating hydrogen oxidation in the HOR process.^{12,34}

Electrochemical HOR performance

The HOR performance of the Ru/Ni–CeO₂ catalyst was assessed using rotating disk electrode (RDE) voltammetry in H₂-saturated 0.1 M KOH electrolyte *via* a standard three-electrode technique, with all electrochemical data being *iR*-corrected. Also, we provided the HOR polarization curves without IR compensation as a comparison (Fig. S6†). First, we investigated the catalytic performance under different reaction conditions and determined that the optimal HOR activity was achieved on Ru/

Ni–CeO₂ with the optimal nickel vacancy annealed at 700 °C (Fig. S7†). Regarding the effect of Ru content, we observed a distinct volcano-like trend in catalytic activity, peaking at a Ru content of 3.75 wt% (Fig. S8 and Table S1†). Suboptimal performance below this level is attributed to insufficient active sites, while excessive Ru content may result in species aggregation, overwhelming the active sites. The optimal Ru content likely promotes a homogeneous distribution of Ru, thereby maximizing active site exposure, which is consistent with previous reports.^{35,36}

Subsequently, the electrocatalytic performances of the synthesized Ni–CeO₂, Ru/Ni, Ru/CeO₂, and Ru/Ni–CeO₂, commercial Ru/C, and benchmark Pt/C were studied under identical conditions for comparison. As observed from the HOR polarization curves (Fig. 4a) of Ru/Ni–CeO₂, the anode current increased rapidly with increasing potential, even exceeding that for the commercial Pt/C, whereas the Ni–CeO₂ showcased a negligible current response, explaining how the EMSI effect between Ru and Ni–CeO₂ markedly accelerated the HOR reaction kinetics. As a comparative experiment, the HOR polarization curves of Ru/Ni–CeO₂ in N₂-saturated electrolyte are documented in Fig. S9,† revealing an almost inconspicuous anode current, confirming that the anode current primarily originates from H₂ oxidation. Fig. 4b presents the HOR polarization curves of Ru/Ni–CeO₂ at speed increments from 400 to 2500 rpm to obtain the relevant kinetic parameters. The Koutecky–Levich curve plotted at an overpotential of 0.1 V reveals a linear relationship between j^{-1} and $\omega^{-1/2}$ with a slope of 4.78 cm² mA⁻¹ rpm^{1/2}, in agreement with the theoretical value of 4.87 cm² mA⁻¹ rpm^{1/2}, proving the two-electron HOR process (inset in Fig. 4b).¹

Furthermore, the kinetic current density (j_k) of the Ru/Ni–CeO₂ catalyst was calculated by measuring the Koutecky–Levich equation (Fig. 4c and e). The j_k of the Ru/Ni–CeO₂ catalyst was as high as 22.2 mA cm⁻² at 50 mV *vs.* RHE, being 2.6-, 2.8-, 4.1-, and 16.8-fold better than those of Ru/CeO₂ (8.3 mA cm⁻²), Ru/Ni (5.4 mA cm⁻²), commercial Pt/C (7.8 mA cm⁻²) and commercial Ru/C (1.3 mA cm⁻²) catalysts, respectively. As displayed in Fig. 4d and e, we acquired the j_0 of the studied electrocatalysts by obtaining the linear fit data in the micropolarization region (–5 to 5 mV). As predicted, the j_0 of Ru/Ni–CeO₂ was calculated to be 3.27 mA cm⁻², and the normalized specific activity (SA) was 0.32 mA cm⁻² (Table S2†), indicating a significant intrinsic activity of Ru/Ni–CeO₂.³⁷ The polarization curves of Ru/Ni–CeO₂ before and after the accelerated durability test (ADT) are shown in Fig. 4f and Fig. S10.† Remarkably, the two HOR polarization curves almost overlap after 1000 cycles, and the limiting current density exhibit only a minimal decrease after 5000 cycles. This performance is notably superior to that of commercial Pt/C. In Fig. S12,† the Ru/Ni–CeO₂-after HOR maintains its initial morphology with only minor structural collapse. Additionally, we observed that Ru, Ce⁴⁺ and Ni⁰ signals were still present in Ru/Ni–CeO₂ after 1000 CVs (Fig. S13†). These results demonstrate that the surface structure of the catalyst was robust enough to protect the active center from alkali attack during the HOR. Moreover,

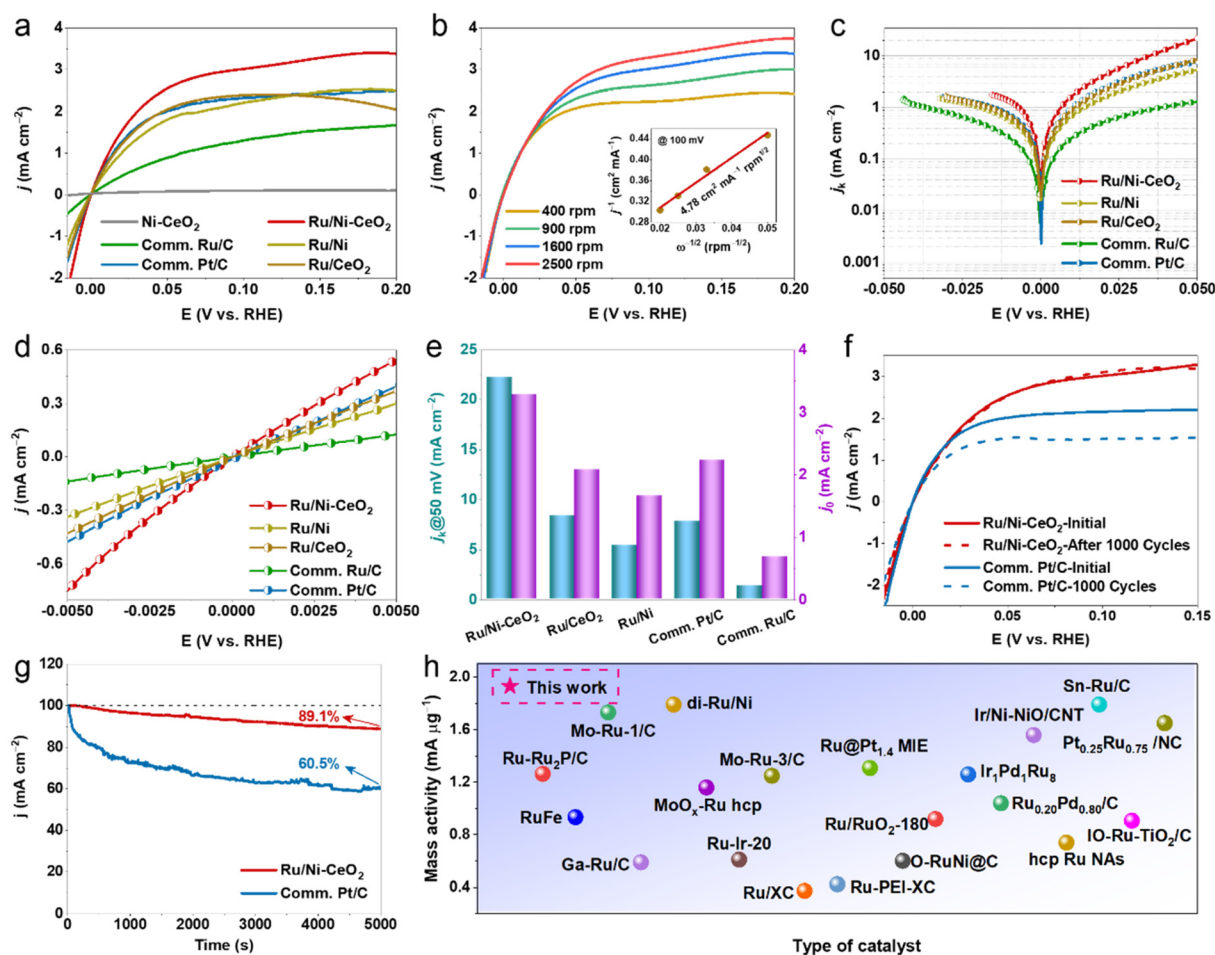


Fig. 4 (a) HOR polarization curves of catalysts in H_2 -saturated 0.1 M KOH at the rotating speed of 1600 rpm. (b) Polarization curves of Ru/Ni–CeO₂ in H_2 -saturated 0.1 M KOH solution at the rotating speeds varied from 400 to 2500 rpm (inset shows corresponding Koutecky–Levich plot). (c) Tafel plots. (d) Linear fitting curves in micropolarization region. (e) Summarization of j_k and j_0 of studied electrocatalysts. (f) HOR polarization curves of Ru/Ni–CeO₂ and commercial Pt/C in H_2 -saturated 0.1 M KOH before and after 1000 CVs. (g) Accelerated durability test and chronoamperometry at 50 mV response of Ru/Ni–CeO₂ and commercial Pt/C. (h) Comparison of the MA with other recently reported alkaline HOR electrocatalysts.

we evaluated the ability of Ru/Ni–CeO₂ to continuously catalyze the HOR using chronoamperometry potentials ($j \approx i$) at 50 mV, which maintained the HOR reactivity throughout the test, indicating good stability (Fig. 4g). Remarkably, the mass activity (MA = 1.93 mA g_{Ru}⁻¹) of Ru/Ni–CeO₂ at 50 mV far exceeded those of most noble metal-based catalysts (Fig. 4h and Table S3†).

Both HBE and OHBE are regarded as active descriptors in the alkaline HOR process.³⁸ Usually, the hydrogen underpotential deposition peak (H_{upd}) is directly associated with HBE on the cyclic voltammetry curve (CV). Lower peaks in the H_{upd} region potentials reflect decreased HBE, which is more advantageous for the HOR process.³⁹ As illustrated in Fig. 5a, the H_{upd} peak potential of Ru/Ni–CeO₂ was more negative than those of Ru/Ni, Ru/CeO₂ and Pt/C, revealing the weaker adsorption of hydrogen on Ru sites. Considering that Ru was easily oxidized in the H adsorption/desorption potential region, it was inappropriate to calculate the ECSA by measuring the charge associated with the H_{upd} region.⁴⁰ According to pre-

vious research, CO stripping is achieved through the reaction to form OH_{ads} and CO_{ads} ($\text{CO}_{\text{ads}} + \text{OH}_{\text{ads}} \rightarrow \text{CO}_2 + \text{H}_2\text{O}$, where CO_{ads} and OH_{ads} represent the radicals adsorbed on the active sites).⁴¹ Consequently, CO-stripping experiments were carried out to identify the ECSA values and to monitor the binding strength of *OH on the investigated catalysts. The ECSA value of Ru/Ni–CeO₂ was 62.1 m² g⁻¹, higher than those of Ru/Ni (30.5 m² g⁻¹), Ru/CeO₂ (19.2 m² g⁻¹) and commercial Pt/C (35.7 m² g⁻¹), as shown in Fig. S11 and Table S2.† The larger ECSA indicated that the more catalytic active sites could adsorb more active intermediates and accelerate the reaction speed.⁴² Meanwhile, as exhibited in Fig. 5b, the CO-stripping peak of Ru/Ni–CeO₂ (0.612 V) was more negative compared to Ru/Ni (0.643 V), Ru/CeO₂ (0.741 V) and commercial Pt/C (0.731 V). This suggested that Ru/Ni–CeO₂ possessed a stronger binding affinity for OH_{ads} .

CO tolerance is a desirable characteristic for hydrogen fuel cell applications due to the primary industrial production of hydrogen coming from the reforming of hydrocarbons by

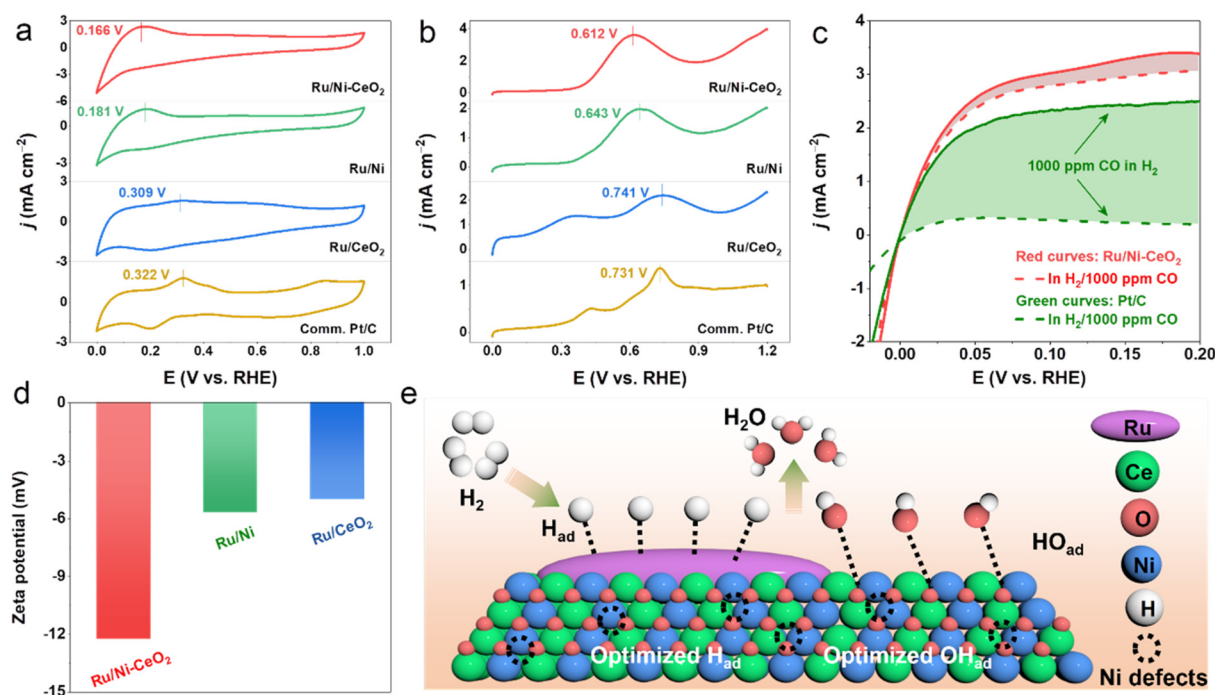


Fig. 5 (a) CV curves of Ru/Ni, Ru/CeO₂, Ru/Ni-CeO₂ and commercial Pt/C in N₂-saturated 0.1 M KOH. (b) CO stripping curves of Ru/Ni, Ru/CeO₂, Ru/Ni-CeO₂ and commercial Pt/C. (c) Comparison of the HOR polarization curves of Ru/Ni-CeO₂ and commercial Pt/C in H₂-saturated 0.1 M KOH containing 1000 ppm CO. (d) Zeta potential of Ru/Ni, Ru/CeO₂, and Ru/Ni-CeO₂. (e) Schematic illustration of HOR catalysis on the Ru/Ni-CeO₂.

natural gas.⁴³ In order to evaluate the CO tolerance of Ru/Ni-CeO₂ catalysts, the electrochemical performance was conducted in H₂-saturated (containing 1000 ppm CO) 0.1 M KOH. As depicted in Fig. 5c, Ru/Ni-CeO₂ maintained excellent HOR activity without an evident decrease, whereas the activity of the commercial Pt/C decreased dramatically, implying the impressive CO tolerance of Ru/Ni-CeO₂.⁴⁴ In addition, the zeta-potential (Fig. 5d) of Ru/Ni-CeO₂ (-12.3 mV) was more negative than those of Ru/Ni (-5.73 mV) and Ru/CeO₂ (-5.05 mV). More negative zeta-potentials imply stronger OHBE, consistent with the results of CO stripping experiments.⁴⁵

Catalytic mechanism

Exploring the mechanism was essential to understanding the excellent HOR performance of Ru/Ni-CeO₂. The CV investigation results (Fig. 5a) showed that the HBE of Ru/Ni-CeO₂ was the weakest, indicating weak H binding to Ru/Ni-CeO₂. LSV tests showed that Ru/Ni-CeO₂ exhibited the best HOR activity, while Ni-CeO₂ was nearly inactive (Fig. 4a), confirming that Ru served as the active site for hydrogen adsorption (H_{ads}). Additionally, the CO-oxidation peak of Ru/Ni-CeO₂ was the most negative, implying that Ru/Ni-CeO₂ had the strongest bonding with OH⁻ based on the correlation between the CO stripping potential and the adsorption strength of OH⁻ (Fig. 5b). The augmented OH_{ads} adsorption behavior and surface structure containing Ni vacancies would boost the trapping of OH_{ads} species on the surface of Ni-CeO₂, thus acceler-

ating the Volmer step in the alkaline HOR process through a bifunctional mechanism at the boundary between the Ru and Ni-CeO₂ components, involving the adsorption of H_{ad} compounds on Ru and that of OH_{ad} on Ni-CeO₂.^{46,47} Consequently, we proposed a bifunctional mechanism for Ru/Ni-CeO₂ by combining CV and CO stripping analysis (Fig. 5e). Specifically, Ru served as a favorable adsorption site for H_{ad}, and H₂ molecules first dissociated and adsorbed on the Ru metal surface to form H_{ad} (denoted as Ru-H_{ad}). Meanwhile, Ni-CeO₂ became an adsorption site for OH to form OH_{ad} (marked as Ni-CeO₂-OH_{ad}), and the last two intermediates (H_{ad} and OH_{ad}) reacted to form H₂O. It is observed that the synergistic adsorption of H and OH⁻ species between Ru and Ni-CeO₂ in the Ni vacancy-rich Ru/Ni-CeO₂ heterogeneity promoted HOR kinetics, thus exhibiting remarkable HOR activity.⁴⁸

Conclusion

Overall, we proved the availability of the Ru/Ni-CeO₂ microsphere structure as a highly efficient electrocatalyst for the hydrogen oxidation reaction (HOR) in alkaline environments. The as-prepared Ru/Ni-CeO₂ showcased admirable HOR performance with a dramatically improved exchange current density and electrochemical surface area, outperforming comparable catalysts and even commercial Pt/C. The remarkable characteristics of Ru/Ni-CeO₂ were ascribed to the EMSI between Ru and Ni-CeO₂ components, resulting in a redistri-

bution of the surface charge density. The charge redistribution was evidenced to optimize the adsorption of hydrogen (H) and hydroxyl (OH) species, leading to enhancement of the HOR kinetics. Furthermore, Ru/Ni–CeO₂ exhibited remarkable long-term stability and CO tolerance. This work offers new insights into advanced catalyst design for the efficient exploitation of hydrogen fuels, benefiting the development of a zero-carbon economy.

Author contributions

Shuqing Zhou: Writing – original draft, methodology. Yi Liu: Methodology, writing – original draft. Lianrui Cheng: Validation, investigation. Tayirjan Taylor Isimjan: Writing – review & editing. Jianniao Tian: Writing – review & editing, supervision. Xiulin Yang: Writing – review & editing, supervision.

Data availability

The data supporting this article have been included as part of the ESI.†

Conflicts of interest

There are no conflicts to declare.

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