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# Comprehensive review of micro/nanostructured ZnSnO<sub>3</sub>: characteristics, synthesis, and diverse applications

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Generally, zinc stannate (ZnSnO<sub>3</sub>) is a fascinating ternary oxide compound, which has attracted significant attention in the field of materials science due to its unique properties such as high sensitivity, large specific area, non-toxic nature, and good compatibility. Furthermore, in terms of both its structure and properties, it is the most appealing category of nanoparticles. The chemical stability of ZnSnO<sub>3</sub> under normal conditions contributes to its applicability in various fields. To date, its potential as a luminescent and photovoltaic material and application in supercapacitors, batteries, solar cells, biosensors, gas sensors, and catalysts have been extensively studied. Additionally, the efficient energy storage capacity of ZnSnO<sub>3</sub> makes it a promising candidate for the development of energy storage systems. This review focuses on the notable progress in the structural features of ZnSnO<sub>3</sub> nanocomposites, including the synthetic processes employed for the fabrication of various ZnSnO<sub>3</sub> nanocomposites, their intrinsic characteristics, and their present-day uses. Specifically, we highlight the recent progress in ZnSnO<sub>3</sub>-based nanomaterials, composites, and doped materials for their utilization in Li-ion batteries, photocatalysis, gas sensors, and energy storage and conversion devices. The further exploration and understanding of the properties of ZnSnO<sub>3</sub> will undoubtedly lead to its broader implementation and contribute to the advancement of next-generation materials and devices.

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## 1. Introduction

The creation and fabrication of materials at the nanoscale have witnessed a significant breakthrough in the 21<sup>st</sup> century, revolutionizing numerous industries, including photocatalysis, energy storage and conversion systems, biosensors, and biological applications. Binary metal-oxides, such as TiO<sub>2</sub>, ZnO, and SnO<sub>2</sub>, which are known for their favourable optical and electrical properties, have been extensively investigated and widely applied in various sectors, such as photovoltaic devices,<sup>1</sup> thin-film displays,<sup>2</sup> electrochromic systems,<sup>3</sup> and gas sensing technologies.<sup>4</sup> Nevertheless, the practical application of these materials is impeded by their poor thermal and chemical stability when exposed to different environments.

Accordingly, to overcome this limitation and improve their characteristics, researchers are actively engaged in the development of ternary oxides, such as ZnSnO<sub>3</sub> and In–Zn–O.<sup>5–7</sup> Among them, nanostructures of ZnSnO<sub>3</sub> (various nano shapes, such as wires, rods, rings, tubes, cubes, and spheres) have attracted considerable interest owing to their advantageous chemical sensitivity, wide energy bandgap, high transmittance percentage,

electron mobility, low price, non-toxicity, and earth abundance.<sup>5,8–13</sup> The performance of energy storage devices and catalysis is greatly affected by the morphology, structure, and physical characteristics of the active electrode materials. Furthermore, LN-type ZnSnO<sub>3</sub>, which possesses a high spontaneous polarization (theoretical value  $P_r \approx 59 \mu\text{C cm}^{-2}$ ), has been experimentally observed in epitaxial thin films with a value of  $P_r \approx 47 \mu\text{C cm}^{-2}$ .<sup>14</sup> These films demonstrate superior photocatalytic activity and exhibit piezoelectric properties.

ZnSnO<sub>3</sub> possesses remarkable morphological properties, making it an attractive material with diverse energy and biological applications. The synthesis and fabrication techniques employed greatly influence the morphological characteristics of ZnSnO<sub>3</sub>. Various factors, such as the capping agent, surfactant, reaction temperature, annealing temperature, concentration of metal precursors, and reaction time, play a crucial role in the development of different synthetic processes.<sup>15–17</sup> Researchers have reported the synthesis of well-organized ZnSnO<sub>3</sub> nanopowders, composites, and films using a range of methods. These methods include solid-state,<sup>18</sup> sol–gel,<sup>19,20</sup> ion-exchange,<sup>21</sup> high temperature calcination,<sup>22,23</sup> thermal evaporation,<sup>24,25</sup> magnetron sputtering,<sup>26–29</sup> hydrothermal process,<sup>8,12</sup> laser ablation,<sup>30</sup> and vapor deposition.<sup>31</sup> The different crystal structures of ZnSnO<sub>3</sub>, such as that with the  $Pm\bar{3}m$ ,  $R\bar{3}$ , and  $R3c$  space groups, have been extensively investigated. The perovskite structure (with the  $Pm\bar{3}m$

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space group) of  $\text{ZnSnO}_3$ , together with its face-centred cubic and orthorhombic phases can be achieved through the calcination of various  $\text{ZnSn(OH)}_6$  precursors. Recent studies have focused on the advancements in  $\text{ZnSnO}_3$  nanostructures for energy and biological applications. However, considering the limited literature to date on this subject, the present review serves as a necessary contribution to bridge the knowledge gap.

This review presents the latest progress in the field of  $\text{ZnSnO}_3$ -based nanomaterials, composites, and doped materials, focusing on their application in the key areas of energy and biology.  $\text{ZnSnO}_3$ -based materials have attracted significant interest due to their potential application in energy storage and conversion technologies, such as lithium/sodium-ion batteries and dye-sensitized solar cells. Additionally, they demonstrate promising prospects as catalysts for the removal of dye/organic pollutants and as gas sensors for various biological uses. This review provides comprehensive insights into the advancements and potential of  $\text{ZnSnO}_3$ -based materials in these specific domains.

## 2. Crystal structure and physical properties

$\text{ZnSnO}_3$  exhibits various types of crystal structures, including perovskite, ilmenite,  $\text{LiNbO}_3$ -type,  $\text{CdSnO}_3$ -type,  $\text{HgSnO}_3$ -type, and post-perovskite with the  $Pm\bar{3}m$ ,  $R\bar{3}$ ,  $R3c$ ,  $Pnma$ ,  $R3c$ , and  $Cmcm$  space groups, respectively, which are all feasible.

$\text{ZnSnO}_3$  with the  $Pm\bar{3}m$ ,  $R\bar{3}$ , and  $R3c$  space groups has been the subject of numerous study. The perovskite structure (with  $Pm\bar{3}m$  space group) of  $\text{ZnSnO}_3$  includes face-centred cubic (FCC) and orthorhombic phases, which is typically synthesized by annealing different  $\text{ZnSn(OH)}_6$  precursors. In terms of surface energy,  $\text{ZnSnO}_3$  crystals with an FCC structure generally exhibit the order of  $(111) < (100) < (110)$ . This demonstrates that the normal surfaces of  $\text{ZnSnO}_3$  crystals tend to exhibit  $\{111\}$  facets, while facets with higher surface energies such as  $\{100\}$  or  $\{110\}$  may not appear during realistic thermodynamic growth processes.<sup>32</sup>

Zinc tin oxide/zinc stannate occur in two individual oxides with distinct crystal structures and varying Zn/Sn ratios. These

oxides are known as orthorhombic or perovskite  $\text{ZnSnO}_3$  and cubic spinel-type  $\text{Zn}_2\text{SnO}_4$ .<sup>33</sup> Based on the available data,  $\text{Zn}_2\text{SnO}_4$  demonstrates higher thermal stability compared to  $\text{ZnSnO}_3$ .  $\text{Zn}_2\text{SnO}_4$  possesses a cubic spinel arrangement, which has been previously established as the most thermodynamically stable form. Alternatively,  $\text{ZnSnO}_3$  is formed under non-equilibrium conditions, such as high pressure, suggesting its thermodynamic metastability as a crystal phase. The formation of metastable  $\text{ZnSnO}_3$  requires high pressure and energy conditions. The phase transition from metastannate to orthostannate begins at approximately 500 °C, with an activation energy of around 0.965 eV. Subsequently, recrystallization occurs, leading to the formation of the orthostannate phase with an inverse spinel structure, which is typically observed at around 750 °C. This investigation provides valuable insights into the behaviour of perovskite  $\text{ZnSnO}_3$  undergoing a phase change to inverse spinel  $\text{Zn}_2\text{SnO}_4$  during calcination.<sup>33</sup>



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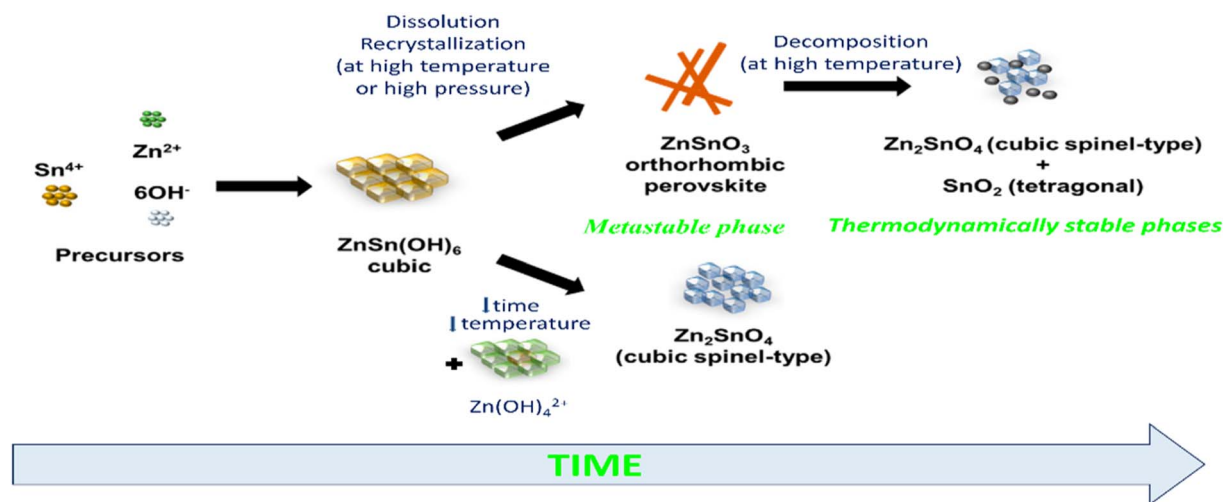


Fig. 1 Schematic illustration of phase rearrangements during the hydrothermal process of zinc stannate nanostructures as a function of reaction energy and synthesis time.<sup>34</sup>

Previous studies have indicated that metastable  $\text{ZnSnO}_3$  structures tend to degrade at temperatures exceeding 500 °C. However, Rovisco *et al.* demonstrated that breakdown can take place at much lower temperatures (e.g., 220 °C for 24 h) or longer reaction durations (e.g., 200 °C for 36 h) because of the high energy involved in the hydrothermal technique.<sup>34</sup> This underscores the benefits of utilizing hydrothermal techniques to acquire metastable nanostructures composed of multiple components, such as  $\text{ZnSnO}_3$ , at reduced temperatures. It also underscores the requirement of carefully managing and comprehending all aspects of fabrication to attain the targeted structures successfully. The growth mechanisms of nanostructures throughout the reaction period pose a significant challenge in the synthesis of nanomaterials. Particularly, in the case of fabricating  $\text{ZnSnO}_3$  nanowire, the metastable nature of this phase adds complexity to its complete comprehension.<sup>34</sup> By increasing the reaction time and overall energy available, the development of nanostructures and their corresponding phases can be observed, which is primarily due to the meticulous optimization of the physio-chemical parameters employed in this study. Their primary objective was to generate  $\text{ZnSnO}_3$  nanowires, but the formation of  $\text{Zn}_2\text{SnO}_4$  nanostructures was also observed, particularly under the conditions of very short synthesis durations, lower temperatures, and smaller reaction volumes. This observation implies that the formation of  $\text{Zn}_2\text{SnO}_4$  requires comparatively less energy. The aforementioned procedure is illustrated in Fig. 1.

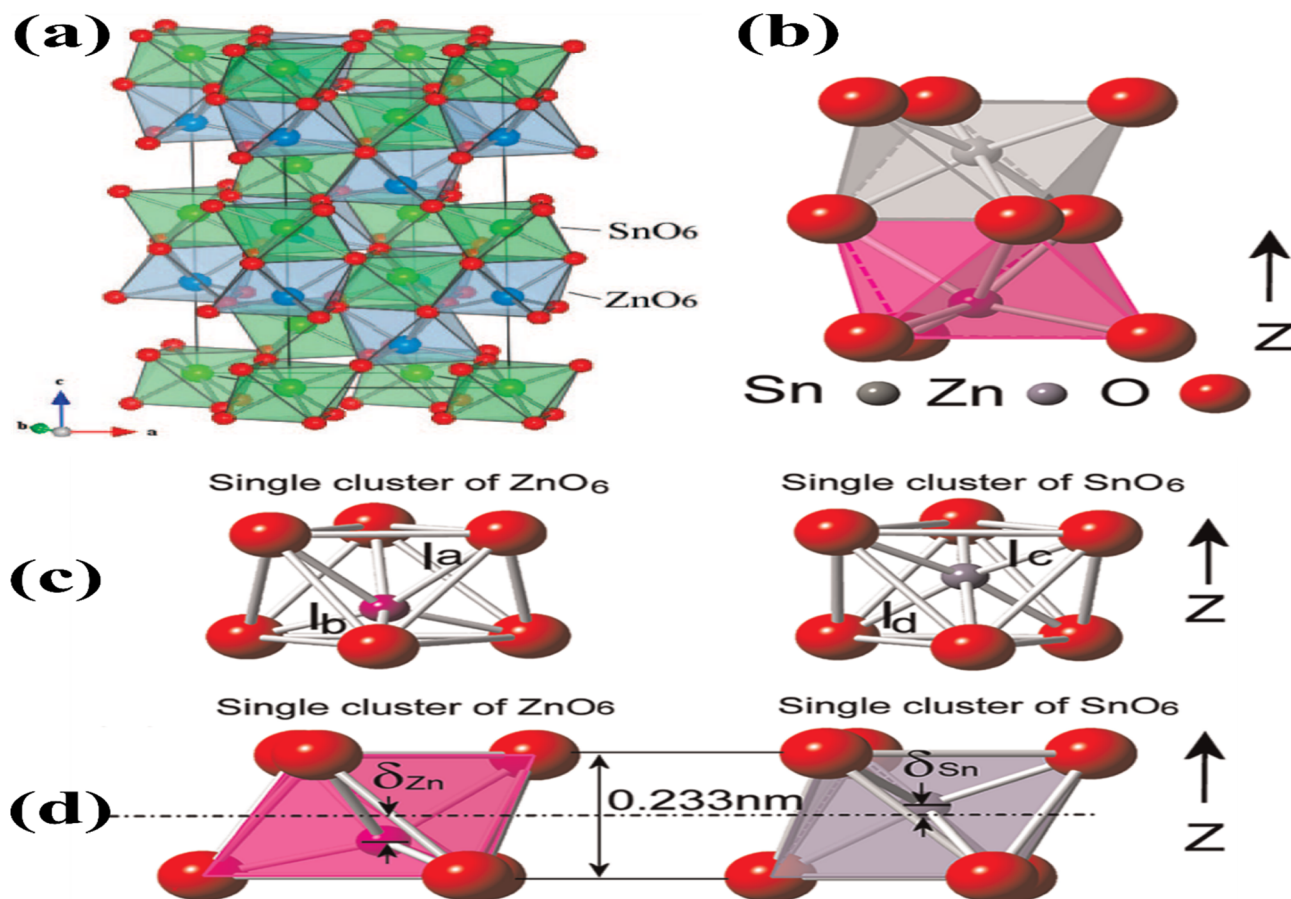
The  $\text{LiNbO}_3$ -type configuration (*R3c*) is distinguished by the substantial displacement of Zn atoms, which is caused by the robust covalent bonds formed between Zn and three oxygen atoms. This bonding configuration gives rise to piezoelectric, ferroelectric, pyroelectric, and non-linear optical properties. The crystal lattice of  $\text{ZnSnO}_3$  in the LN-type arrangement, as depicted in Fig. 2a, is comprised of interconnected octahedral units. Interlocking occurs between the Zn octahedra, where each octahedron shares its corners with another octahedron of

the same type. Similarly, Sn octahedra exhibit corner sharing, forming connections with other Sn octahedral structures. The cation arrangement follows a pattern of Sn–Zn–vacancy–Sn–Zn–vacancy–Sn, aligned along the z-axis. In  $\text{ZnSnO}_3$ , the bond valence sums for Zn, Sn, and O were calculated to be 1.79, 4.08, and 1.96, respectively. Notably, the Zn–O bond in  $\text{ZnSnO}_3$  is found to be under-bonded compared to the ideal values. The Sn–O distances in  $\text{ZnSnO}_3$  were observed to be 0.2005 nm ( $\times 3$ ) and 0.2094 nm ( $\times 3$ ), deviating from the distances typically observed in perovskite-type stannates, which feature  $\text{SnO}_6$  octahedra.<sup>35</sup>

Fig. 2b illustrates the presence of two octahedral structures, where one is composed of  $\text{ZnO}_6$  and the other  $\text{SnO}_6$ . The  $\text{SnO}_6$  and  $\text{ZnO}_6$  octahedra are connected to the neighbouring octahedra through shared edges and faces. Fig. 2c illustrates a single  $\text{ZnO}_6$  and  $\text{SnO}_6$  cluster. The Zn–O bonding lengths consist of three long bonds (approximately 0.2308 nm) on the upper side and three short bonds (approximately 0.2040 nm) on the bottom side, which are labelled as la and lb in the  $\text{ZnO}_6$  cluster, respectively. The Sn–O bonding lengths in the  $\text{SnO}_6$  cluster, labelled as lc and ld, consist of three short bonds (approximately 0.2008 nm) on the upper side and three long bonds (approximately 0.2093 nm) on the lower side. Fig. 2d shows individual clusters of  $\text{ZnO}_6$  and  $\text{SnO}_6$ , highlighting the variation in bonding length along the z-axis. In the  $\text{ZnO}_6$  cluster, the Zn ion exhibits a displacement ( $\delta_{\text{Zn}}$ ) of 0.5 Å, while in the  $\text{SnO}_6$  cluster, the Sn ion has a displacement ( $\delta_{\text{Sn}}$ ) of 0.2 Å. Along the z-axis, the Zn ion experiences a larger displacement than the Sn ion, resulting in the creation of spontaneous polarization, which is the origin of piezoelectricity in this material.<sup>36,37</sup>

In ref. 38 and 39, density functional theory (DFT) and the extended gradient estimation were employed to investigate the structural, electrical, and optical properties of  $\text{ZnSnO}_3$ . The analysis of the electronic structures revealed that  $\text{ZnSnO}_3$  is a semiconductor characterized by a direct band gap of 1.0 eV. The examination of the optical spectra revealed that inter-band





**Fig. 2** (a) Crystal structure of ZnSnO<sub>3</sub> is visualized in a three-dimensional image, revealing the arrangement of atoms in an octahedral framework [reprinted with permission from *J. Am. Chem. Soc.* 2008, **130**, 21, 6704–6705. Copyright 2008, the American Chemical Society]. (b) Lower side of the structure consists of ZnO<sub>6</sub> octahedra, while the upper side is formed by SnO<sub>6</sub> octahedra. (c) Separate and distinct clusters of ZnO<sub>6</sub> and SnO<sub>6</sub>. (d) Displacement of Zn ions ( $\delta_{Zn}$ ) and Sn ions ( $\delta_{Sn}$ ) from their equilibrium positions along the z-axis causes a variation in the bonding length between oxygen (O) and Zn (O–Zn–O) or Sn (O–Sn–O) atoms, respectively.<sup>48</sup>

transitions occur in the ZnSnO<sub>3</sub> compound. These transitions occur between the O 2p levels in the valence band (VB) and either the Sn 5s level or the higher CB Zn 3d levels in the lower energy level. Additionally, inter-band transitions were observed between the O 2p levels and either the Sn 5p or Zn 4p conduction bands (CB) in the higher energy level. These transitions contributed to the computed optical spectra.

Dielectric properties are also crucial properties for the application of nanoparticles as dielectric materials.<sup>40–42</sup> ZnSnO<sub>3</sub> materials display excellent electromagnetic wave attenuation characteristics and a wide frequency range, making them suitable for various applications such as ground-penetrating radar systems, microwave absorbers, communication systems, and energy storage devices.<sup>43–45</sup> To obtain the real ( $\epsilon_1$ ) and imaginary ( $\epsilon_2$ ) part of the dielectric constant, the Kramers–Kronig equation was utilized.<sup>46</sup> Given that ZnSnO<sub>3</sub> possesses a hexagonal shape, the evaluation was focused on incoming light polarized along the [1 0 0] and [0 0 1] crystallographic axes. They observed that there was no significant anisotropy in both the real and imaginary parts of the equation. The peaks in  $\epsilon_2$  were associated with electron excitation. Furthermore, the computed static dielectric

constant,  $\epsilon_1(0)$ , of ZnSnO<sub>3</sub> along the [1 0 0] and [0 0 1] directions was determined to be 4.05 and 3.96 eV, respectively.<sup>39</sup> These results are significantly lower than that of BaTiO<sub>3</sub> (5.12) and PbZrO<sub>3</sub> (5.34), indicating the distinct dielectric behaviour of ZnSnO<sub>3</sub> compared to these materials.<sup>47</sup>

ZnSnO<sub>3</sub> also exhibits superior ferroelectric properties. Shin *et al.* conducted research on the ferroelectric characteristics of ZnSnO<sub>3</sub>.<sup>49</sup> They examined the hysteresis loop of a Pt/ZnSnO<sub>3</sub>/SrRuO<sub>3</sub> capacitor at a measurement frequency of 10 kHz to investigate its ferroelectric properties. This exhibited a coercive electric field of 130 kV cm<sup>−1</sup> and improved remnant polarization of 47 C cm<sup>−2</sup> ( $2P_r$  of 94 C cm<sup>−2</sup>). The epitaxial ZnSnO<sub>3</sub> demonstrated a saturation polarization of 58 C cm<sup>−2</sup>, which was marginally higher than the residual polarization. This observation suggests the presence of a well-formed crystalline structure in the material. To gain a deeper understanding, they investigated the hysteresis loops across a range of frequencies. This analysis aimed to examine the behaviour of the capacitor at various frequency regimes. As depicted in Fig. 3a, the coercivity increased gradually as the measurement frequency increased. Additionally, a square pulse with a voltage of 5 V was employed



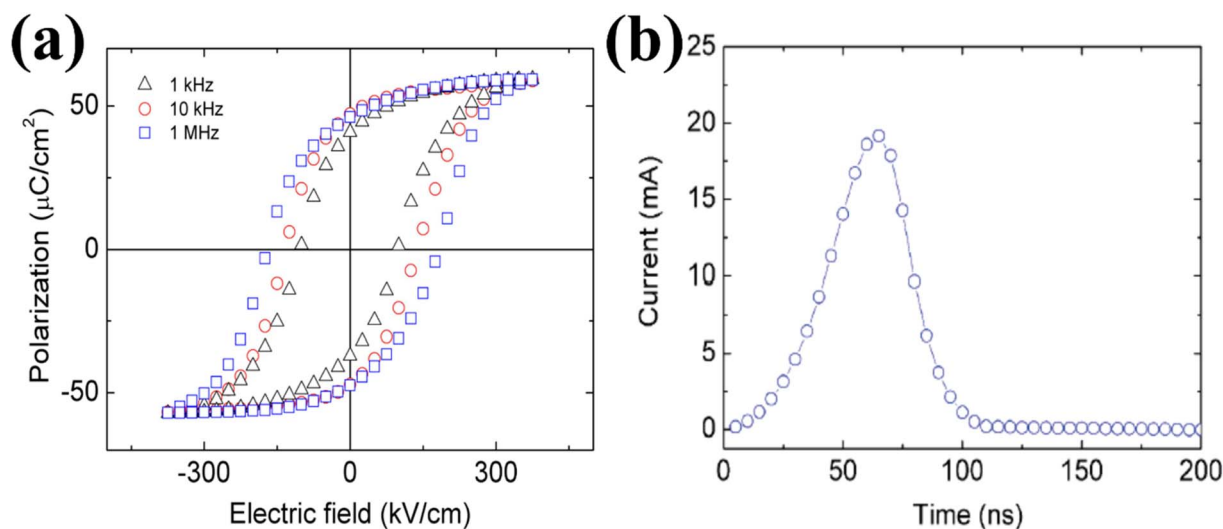


Fig. 3 (a) Hysteresis loops of the Pt/ZnSnO<sub>3</sub>/SrRuO<sub>3</sub> capacitor observed at different measurement frequencies. (b) Switching-current as a bias of 5 V. The fast-switching behaviour is within 100 ns.<sup>49</sup>

to evaluate the switching current. The switching behaviour was found to be rapid, with a fast-switching time of 100 ns, as shown in Fig. 3b. This indicates that the capacitor can transition between states efficiently and quickly.

To model the interactions between ions and electrons, projector expanded wave potentials were employed, considering the effects of both ions and electrons.<sup>49</sup> The exchange and correlation energies of the electrons were calculated using a local density approximation, providing insights into their behaviour in the system. The estimated lattice constants were  $a = 5.24 \text{ \AA}$  and  $c = 13.88 \text{ \AA}$  per formula unit cell, with a cell volume of  $54.92 \text{ \AA}^3$ . Fig. 4 presents the measurements of the deviations of Zn and Sn ions from the oxygen octahedral core. The data revealed that the rearrangement of the A cation (Zn) was  $0.55 \text{ \AA}$ , which exceeds the rearrangement of the B cation (Sn) at  $0.21 \text{ \AA}$ . This disparity can be attributed to the larger available space in the A site, facilitating greater movement for the Zn atom. The difference in rearrangement arises from the fact that Zn, which has a covalent radius of  $1.31 \text{ \AA}$ , is smaller in size compared to Sn, with a covalent radius of  $1.41 \text{ \AA}$ .<sup>49</sup> Employing the Berry phase technique, a polarization value of  $60 \text{ C cm}^{-2}$  was determined along the pseudo cubic<sup>109</sup> direction. This value closely matches both their experimental findings and the results obtained from analytical measurements based on the ionic rearrangements and atomic valences.

The preferred ground-state structure of ZnSnO<sub>3</sub> was found to be the LN-type (LiNbO<sub>3</sub>-type) phase, given that it exhibits a lower total energy compared to the IL-type (ilmenite-type) structure. The energy difference between the LN and IL phases is merely  $0.09 \text{ eV}$  per unit of formula, indicating the potential occurrence of a structural shift from the LN to the IL when exposed to high pressure or temperature. However, investigations by Gao *et al.* revealed that the total energy-volume curves of the LN and IL phases do not cross when

subjected to compression, indicating that the structural transition is unlikely to occur under high pressure.<sup>39</sup>

Fig. 5a illustrates the calculated  $E$ - $V$  curves for different structures. When  $T = 0 \text{ K}$ , the LN structure exhibits the lowest total energy among the feasible structures. Alternatively, the cubic perovskite structure has a considerably higher total energy, with a difference of up to  $3.76 \text{ eV}$  per formula unit compared to the LN-type structure. This substantial energy variation implies that the formation of the cubic perovskite phase of ZnSnO<sub>3</sub> is challenging to achieve under normal conditions. Additionally, the combined energy of the CdSnO<sub>3</sub>-type phase crosses paths with both the IL and LN phases, indicating the possibility of structural changes in severe settings.<sup>50–52</sup>

Moreover, Fig. 5b illustrates the evaluation of the total energy variation in ZnXO<sub>3</sub> compounds (where, X = Si, Ge, Sn, and Pb) between the IL and LN structures. The results indicate a monotonic decrease in the total energy difference moving from Si ( $0.60 \text{ eV}$ ) to Ge ( $0.37 \text{ eV}$ ) to Sn ( $0.09 \text{ eV}$ ) phases, proposing that the IL phase is strenuously more favourable than the LN phase for these elements. However, for the Pb-containing phase, the LN phase is more energetically favourable than the IL phase. In the range where the zero-point energy (ZPE) correction becomes significant for assessing the relative structural stability, the overall energy gap between ZnSnO<sub>3</sub> in ionic liquid (IL) and layered perovskite (LN) forms is comparatively insignificant. The ZPE calculated from the partial density of states (PDOS) was found to be  $0.23 \text{ eV}$  per unit for the IL-type structure and  $0.31 \text{ eV}$  per unit for the LN structure. When accounting for ZPE, the overall energy gap between the IL and LN structures is approximately  $0.005 \text{ eV}$  per unit. This implies that there is a possibility for these two phases to exist together under normal environmental conditions.<sup>50,53–55</sup>

The lattice parameters of a semiconductor are typically affected by multiple parameters, as follows:<sup>56–59</sup> (i) the concentration of free electrons, which affects the deformation



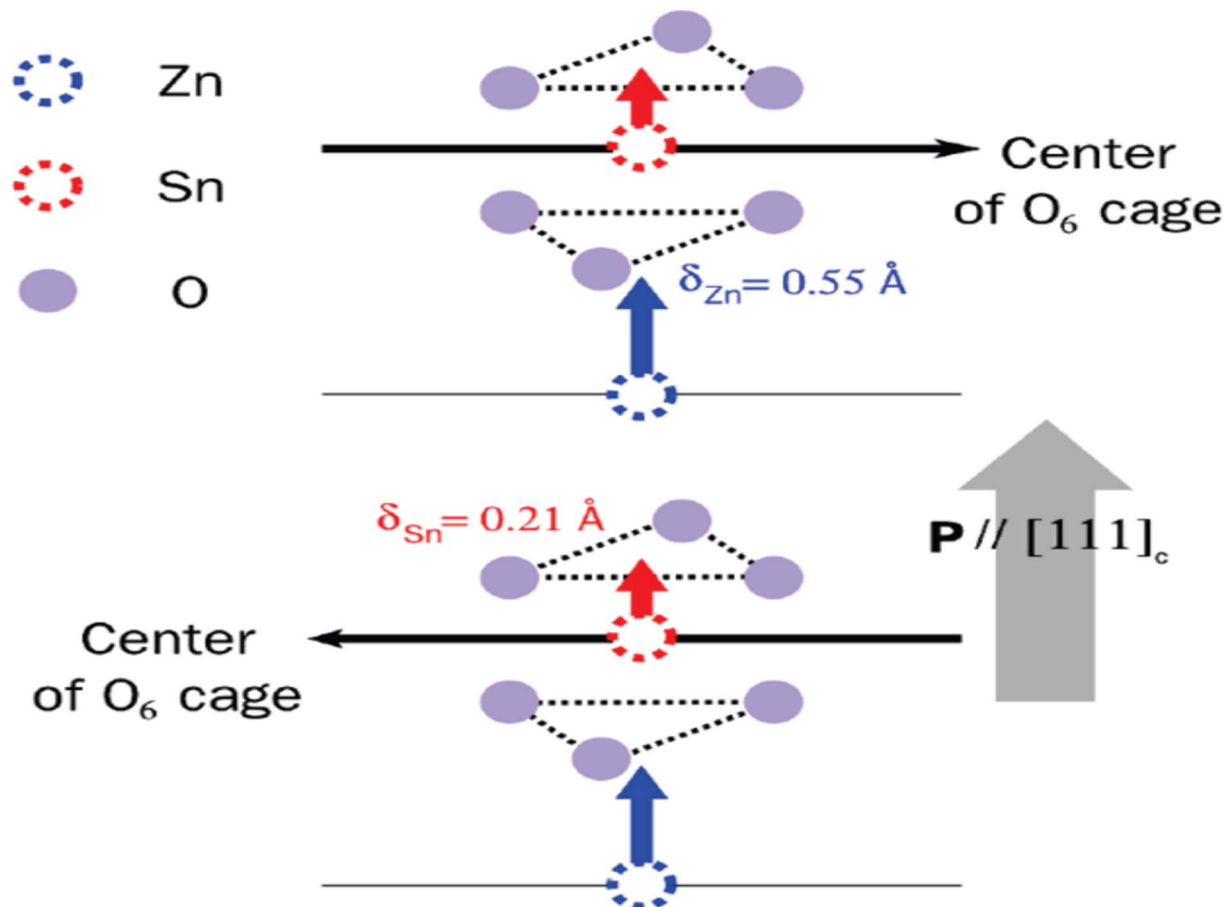


Fig. 4 Ionic rearrangements of Zn and Sn in the R3c phase of ZnSnO<sub>3</sub>.<sup>49</sup>

potential of the conduction-band minimum occupied by these electrons, (ii) the concentration of foreign atoms and defects and their difference in ionic radius compared to the host matrix ion, (iii) external strains caused by factors such as substrate-induced stress, and (iv) temperature. To accurately determine the lattice parameters of a crystalline material, high-resolution X-ray diffraction (HRXRD) is commonly employed. The bond method is utilized with a combination of symmetrical and asymmetrical reflections to measure and analyse the lattice parameters.<sup>60–63</sup> In the case of ZnSnO<sub>3</sub>, the calculated lattice parameters are provided in Table 1.

In ref. 50, the formation enthalpy [ $\Delta H = E_{\text{total}}(\text{ZnSnO}_3) - \{E_{\text{total}}(\text{ZnO}) + E_{\text{total}}(\text{SnO}_2)\}$ ] was estimated to gain insights into the impact of different synthesis pathways in experiments. The calculated formation enthalpies for all the proposed ZnSnO<sub>3</sub> phases are positive, as shown in Table 1. This indicates that ZnSnO<sub>3</sub> is not energetically favourable and cannot be produced through solid-state fabrication pathways such as combining ZnO and SnO<sub>2</sub> under normal environmental conditions. However, experimental evidence suggests that these polymorphs have the potential to remain stable when subjected to extreme conditions, such as elevated pressure and temperature. To comprehend the structural transition of ZnSnO<sub>3</sub>, the enthalpy variation between heterogeneous component oxides such as (ZnO + SnO<sub>2</sub>), ((Zn<sub>2</sub>SnO<sub>4</sub> + SnO<sub>2</sub>)/2), IL- and LN-type phases was also estimated.

This analysis is depicted in Fig. 5c and d). Below 5.9 GPa, the heterogeneous component oxides (h-ZnO + SnO<sub>2</sub>) are more favourable than the various potential phases. In the range of 5.9 to 7.1 GPa, the heterogeneous component oxides of Zn<sub>2</sub>SnO<sub>4</sub> + SnO<sub>2</sub> become increasingly favourable, aligning with experimental results at intermediate pressure levels. At low temperatures and >7.1 GPa, the LN-type ZnSnO<sub>3</sub> phase is more favourable than its constituent phases, which is consistent with experimental data, suggesting the formation of LN-type ZnSnO<sub>3</sub> at 7 GPa. At a pressure of 34.5 GPa, the LN-type phase undergoes a transition to the orthorhombic CdSnO<sub>3</sub>-type phase, which is significantly higher than the transition pressures observed for ZnGeO<sub>3</sub> (15.6 GPa) and MgGeO<sub>3</sub> (17.9 GPa).<sup>51,57,64</sup>

Elastic constants are significant parameters that provide insight into the crystallite structure and bonding strength among atoms. In hexagonal structures such as ZnSnO<sub>3</sub>, the elastic parameters exhibit positive values and adhere to the stability requirement outlined by Born–Huang, suggesting the elastic stability of both the LN and IL phases. Table 2 presents the elastic constants for various space groups.

The structural properties of ZnSnO<sub>3</sub> can be understood based on density functional theory (DFT) calculations. The ZnSnO<sub>3</sub> supercells used in the study consist of 60 atoms (Zn<sub>12</sub>Sn<sub>12</sub>O<sub>36</sub>). The VB of ZnSnO<sub>3</sub> primarily consists of Zn 3d<sup>10</sup>4s<sup>2</sup> states, Sn 5s<sup>2</sup>5p<sup>2</sup> states, and O 2s<sup>2</sup>p<sup>4</sup> states. Fig. 6a and



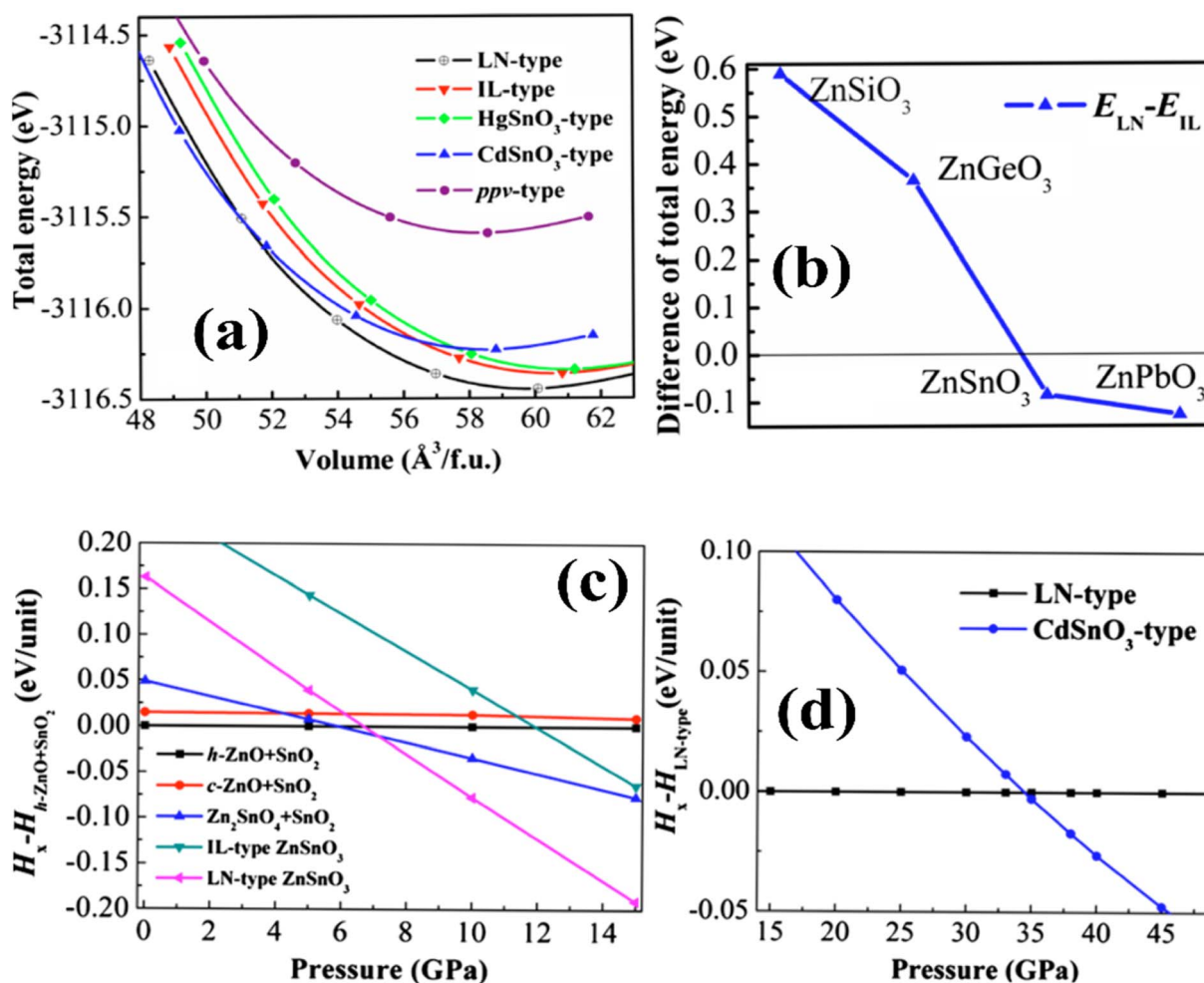


Fig. 5 (a)  $E$ - $V$  curve for six feasible phases. (b) Energy variation in  $\text{ZnXO}_3$ . (c) Enthalpy variation among heterogenous component oxides. (d) Enthalpy variation in LN-type and  $\text{CdSnO}_3$ -type phases.<sup>65</sup>

Table 1 Previously reported lattice constants, total energy ( $\Delta E$ ), and heat of formation ( $\Delta H_f$ ) of  $\text{ZnSnO}_3$  for various space groups

| Space group  | Lattice constants    |                      |                      | $\Delta E$ (eV/f.u.) | $\Delta H_f$ (eV/f.u.) | Ref.      |
|--------------|----------------------|----------------------|----------------------|----------------------|------------------------|-----------|
|              | $a$ ( $\text{\AA}$ ) | $b$ ( $\text{\AA}$ ) | $c$ ( $\text{\AA}$ ) |                      |                        |           |
| $R3c$        | 5.387(5.262)         | 14.344(14.003)       | —                    | 0                    | 0.15                   | 35 and 50 |
| $R\bar{3}$   | 5.419(5.284)         | 14.384(14.091)       | —                    | 0.09                 | 0.23                   |           |
| $Pm\bar{3}m$ | 4.086                | —                    | —                    | 3.76                 | 3.90                   |           |
| $R\bar{3}c$  | 3.429                | 14.387               | —                    | 0.11                 | 0.25                   |           |
| $Pnma$       | 5.422                | 7.994                | 5.428                | 0.22                 | 0.36                   |           |
| $Cmcm$       | 3.082                | 9.934                | 7.653                | 0.85                 | 1.00                   |           |

b illustrate the supercells and the determined density of states (DOS) and partial density of states (PDOS) of  $\text{ZnSnO}_3$ .<sup>67</sup>

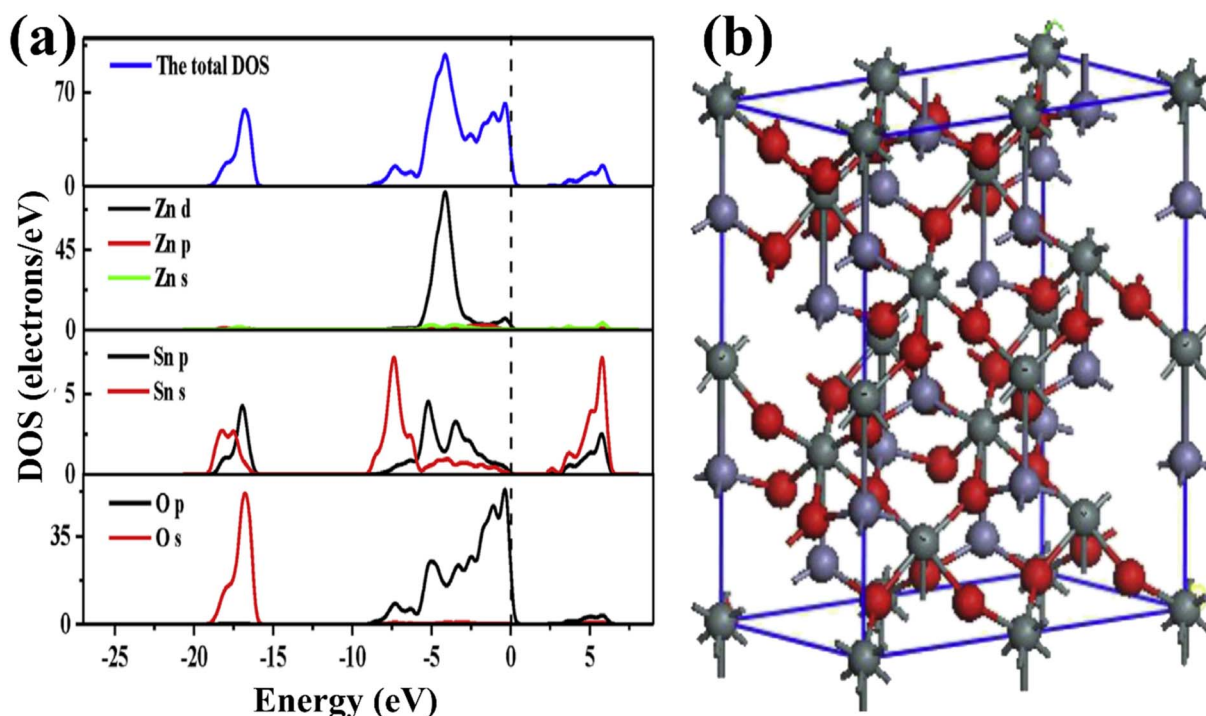
The energy band at  $[-18.4, -15.3]$  eV in the VB is largely filled by O 2s states interspersed with few Sn 4d levels. The energy range of approximately  $[-8.1, -5.5]$  eV displays bands that can be attributed to the presence of Sn 5s states, together with a combination of O 2p and Zn 3d states. On the other hand, within the range of  $[-5.5, -3.4]$  eV, the dominating factor is the Zn 3d states, accompanied by O 2p states and a small

contribution from Sn 5p states. At the upper end of the valence band  $[[3.4, 0]$  eV), the O 2p state takes precedence and dominates. However, it is accompanied by a mixture of Zn 3d and minor presence of Sn 4d states, suggesting substantial hybridization between these states. The Sn 5s and Zn 4s states dominate the CB area of  $[1, 5.5]$  eV, but Sn 5p dominates the Zn 4p state when the energy is more than 5.5 eV. For energies less than 5.5 eV, the Zn PDOS is minimal, and thus negligible. Consequently, the Sn 5s and O 2p states should control the electrical



Table 2 Elastic constants of ZnSnO<sub>3</sub> for various space groups

| Space group | C <sub>11</sub> | C <sub>12</sub> | C <sub>13</sub> | C <sub>14</sub> | C <sub>15</sub> | C <sub>22</sub> | C <sub>23</sub> | C <sub>33</sub> | C <sub>44</sub> | Ref. |
|-------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|-----------------|------|
| <i>R3c</i>  | 310.4           | 137.1           | 100.8           | 1.4             | —               | —               | —               | 175.3           | 77.3            | 66   |
| <i>R3̄</i>  | 268.8           | 132.8           | 78.8            | 14.7            | 14.3            | —               | —               | 185.9           | 2.6             |      |
| <i>Pm3m</i> | 284.4           | 74.1            | —               | —               | —               | —               | —               | —               | 33.8            |      |
| <i>Pnma</i> | 290.8           | 103.8           | 116.7           | —               | —               | 247.9           | 100.4           | 273.3           | 74.7            |      |
| <i>Cmcm</i> | 316.9           | 51.8            | 47.8            | —               | —               | 210.4           | 77.5            | 223.9           | 239.5           |      |

Fig. 6 (a) DOS and PDOS for ZnSnO<sub>3</sub>. (b) Supercells of Zn<sub>12</sub>Sn<sub>12</sub>O<sub>36</sub>.<sup>69</sup>

characteristics of ZnSnO<sub>3</sub>. Both the top of the VB and the bottom of the CB are located at the  $\Gamma$ -point. Consequently, a direct band gap of 1.0 eV is generated, showing that ZnSnO<sub>3</sub> is a semiconductor. Due to the absence of an experimental band gap result for comparison, it is important to note that DFT often underemphasizes the energy gap of semiconductor solids.<sup>67,68</sup>

### 3. Synthesis technique/routes

Over the past few decades, several synthetic techniques have been employed to synthesize micro/nanostructures of ZnSnO<sub>3</sub>. However, to achieve the controlled synthesis of these structures and expand their practical applications, it is crucial to comprehensively summarize the emerging growth mechanisms and develop new techniques. The two primary methods used for the synthesis of nanostructures are solution-based and vapor phase approaches. The size, crystal phase, and crystallinity of the synthesized materials significantly impact their band-gap energies and the separation of charge carriers in semiconductor oxides. Therefore, the quality of the synthetic conditions plays a key function in determining the efficiency of

semiconductor oxides. Comparing with various synthesis routes, the solution-based procedure emerges as a simple and low energy-consuming technique for the production of ZnSnO<sub>3</sub> nanomaterials. By modifying the experimental factors such as solvent, precursor ingredients, and reaction environments, precise control of the nanostructure morphologies can be achieved. This straightforward technique also enables greater control of the particle size of nanostructures. The solution-based methods employed for the synthesis of ZnSnO<sub>3</sub> nanostructures include hydrothermal,<sup>8,12</sup> sol-gel,<sup>19,20</sup> precipitation,<sup>70</sup> electrochemical-deposition,<sup>36</sup> solvothermal,<sup>71</sup> microwave,<sup>72</sup> wet chemical, and electrospinning method.<sup>73</sup> Among the different methods available, the simple sol-gel process is particularly appealing for the synthesis of ZnSnO<sub>3</sub> nanocomposites. This method offers several advantages, including low cost, reliability, good repeatability, and the ability to easily control the physical properties and morphologies.<sup>32</sup>

The activity of ZnSnO<sub>3</sub> can be enhanced through the precise control of various factors, including band gap, size, morphology, crystal structure, surface area, stability, reusability, and preparation of composite materials. Among them,



the crystal structure and specific surface area play a crucial role in improving the performance of ZnSnO<sub>3</sub>. A significant parameter for achieving superior results is a large surface area, which is closely interconnected with the structure and mean particle size of the material. In the study by Yu *et al.*, they synthesized homogeneous ZnSnO<sub>3</sub> nanocubes using a low-temperature solution route.<sup>74</sup> They observed that decreasing the reaction temperature from 80 °C to 0 °C led to a decrease in the particle size of the fabricated ZnSnO<sub>3</sub> nanocubes from 600 nm to 40 nm, while the surface area increased from 36.4 to 109.5 m<sup>2</sup> g<sup>-1</sup>. The reduction in particle size resulted in an increase in the overall surface area of the material. A larger surface area provides more reactive sites, leading to higher surface reactivity, greater surface-to-volume ratios, and increased availability of surface-active sites.

Moreover, vapor phase techniques have also been used for the synthesis of nanostructured materials. These techniques include laser ablation,<sup>75</sup> vapour-liquid-solid,<sup>37</sup> thermal evaporation,<sup>76</sup> molecular beam epitaxy (MBE),<sup>77</sup> metal-organic chemical vapour deposition<sup>78</sup> and magnetron sputtering.<sup>79</sup> The choice of the method for the synthesis of ZnSnO<sub>3</sub> primarily relies on the desired dimensions of the nanostructures. The impact of various fabrication techniques and starting materials

on the morphologies of ZnSnO<sub>3</sub> is presented in Table 3. Furthermore, in the application domain, the influence of the morphology of ZnSnO<sub>3</sub> on its efficiency has also been investigated.

### 3.1. Hydrothermal technique

Hydrothermal synthesis is a widely employed technique for the preparation of ZnSnO<sub>3</sub>, utilizing a solution-based reaction approach. This method involves placing starting material dispersion in a sealed stainless-steel Teflon-lined autoclave. Subsequently, the autoclave is subjected to heating in an oven under precise temperature, duration, and autogenous pressure settings. The temperature typically ranges from 140 °C to 200 °C for the hydrothermal synthesis of zinc stannate.<sup>80</sup> The hydrothermal synthesis of ZnSnO<sub>3</sub> nanoparticles involves the use of different mineralizers such as ammonium hydroxide, sodium hydroxide, hexadecyl trimethyl ammonium bromide, various amines, and sodium carbonate. This leads to the creation of nanocrystals with different shapes (cubic, spherical, and rod like). Careful control of the chemical properties of the mineralizer is essential given that it influences the surface charges of the resulting metal oxides, which is critical

Table 3 ZnSnO<sub>3</sub> micro/nanostructure synthesis processes and morphologies

| Material                | Morphologies                  | Precursors                                                                                                                                                                                  | Method                    | Ref.            |
|-------------------------|-------------------------------|---------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------------|---------------------------|-----------------|
| ZnSnO <sub>3</sub>      | Polyhedral                    | Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O, CTAB/SDBS                                                                                                                                           | Solution-based            | 32              |
| ZnSnO <sub>3</sub>      | Nanocubes                     | SnCl <sub>4</sub> ·5H <sub>2</sub> O/Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O                                                                                                                 | HNO <sub>3</sub> etching  | 92              |
| ZnSnO <sub>3</sub>      | Nanocubes                     | SnCl <sub>4</sub> ·5H <sub>2</sub> O/ZnSO <sub>4</sub>                                                                                                                                      | Solution-based            | 74              |
| ZnSnO <sub>3</sub>      | Hollow-cubes                  | Zn(NO <sub>3</sub> ) <sub>2</sub> /SnCl <sub>4</sub>                                                                                                                                        | HCl etching               | 131             |
| ZnSnO <sub>3</sub>      | Hollow-cages                  | SnCl <sub>2</sub> /Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O                                                                                                                                   | Solution-based            | 132             |
| ZnSnO <sub>3</sub>      | Hollow-cubes                  | SnCl <sub>4</sub> ·5H <sub>2</sub> O/ZnCl <sub>2</sub>                                                                                                                                      | Solution-based            | 125             |
| ZnSnO <sub>3</sub>      | Hollow-cages                  | SnCl <sub>4</sub> ·5H <sub>2</sub> O/Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O                                                                                                                 | Solution-based            | 127             |
| ZnSnO <sub>3</sub>      | Face-centred trigonal         | Na <sub>2</sub> SnO <sub>3</sub> ·3H <sub>2</sub> O/Zn(NO <sub>3</sub> ) <sub>2</sub> ·6H <sub>2</sub> O                                                                                    | Sol-gel                   | 99              |
| ZnSnO <sub>3</sub>      | Polycrystalline ilmenite-type | Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O/SnCl <sub>4</sub> ·5H <sub>2</sub> O/ethylene glycol                                                                                                 | Sol-gel                   | 100             |
| ZnSnO <sub>3</sub>      | Nanostructures                | ZnCl <sub>2</sub> /SnCl <sub>2</sub> /Tepa/trimesic acid                                                                                                                                    | Sol-gel                   | 103             |
| ZnSnO <sub>3</sub>      | Hollow cubes                  | SnCl <sub>4</sub> /ZnCl <sub>2</sub>                                                                                                                                                        | NaOH etching              | 133 and 134     |
| ZnSnO <sub>3</sub>      | Nanoplates                    | Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O, SnCl <sub>4</sub> ·5H <sub>2</sub> O, ethanol, MEA                                                                                                  | Sol-gel                   | 135             |
| ZnSnO <sub>3</sub>      | Hierarchical nanocages        | SnCl <sub>4</sub> ·5H <sub>2</sub> O/Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O/NaOH, (CH <sub>2</sub> ) <sub>6</sub> N <sub>4</sub>                                                            | Hydrothermal              | 128             |
| ZnSnO <sub>3</sub>      | Hollow microspheres           | SnCl <sub>4</sub> ·5H <sub>2</sub> O/Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O/NaOH, CTAB                                                                                                      | Hydrothermal              | 126             |
| ZnSnO <sub>3</sub>      | Nanocages                     | SnCl <sub>4</sub> ·5H <sub>2</sub> O/Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O/NaOH, (CH <sub>2</sub> ) <sub>6</sub> N <sub>4</sub>                                                            | Hydrothermal              | 136             |
| ZnSnO <sub>3</sub>      | Nanocubes                     | Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O/SnCl <sub>4</sub> ·5H <sub>2</sub> O, NaOH                                                                                                           | Hydrothermal              | 137             |
| ZnSnO <sub>3</sub>      | Nanowires                     | ZnCl <sub>2</sub> /SnCl <sub>4</sub> ·5H <sub>2</sub> O/PEG (200)                                                                                                                           | Hydrothermal              | 138             |
| ZnSnO <sub>3</sub>      | Nanowires                     | Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O/SnCl <sub>4</sub> ·5H <sub>2</sub> O/NaOH, PEG (4 K)                                                                                                 | Single-step hydrothermal  | 91              |
| ZnSnO <sub>3</sub>      | Nanowires                     | ZnO/SnCl <sub>4</sub> ·5H <sub>2</sub> O/PEG (10 K)                                                                                                                                         | Hydrothermal              | 139             |
| ZnSnO <sub>3</sub>      | Nanocubes                     | SnK <sub>2</sub> O <sub>3</sub> ·3H <sub>2</sub> O/ZnC <sub>4</sub> H <sub>6</sub> O <sub>4</sub> ·2H <sub>2</sub> O/urea                                                                   | MA hydrothermal           | 72              |
| ZnSnO <sub>3</sub>      | Nanocubes                     | ZnSO <sub>4</sub> ·7H <sub>2</sub> O/Na <sub>2</sub> SnO <sub>3</sub> ·3H <sub>2</sub> O                                                                                                    | Ionic substitution        | 140             |
| ZnSnO <sub>3</sub>      | Nanobelts/microbelts          | Zn and Sn powder/graphite powder                                                                                                                                                            | Vapour-liquid-solid       | 36              |
| ZnSnO <sub>3</sub>      | Triangular-belts              | Zn and Sn/graphite powder                                                                                                                                                                   | Vapour-liquid-solid       | 37              |
| ZnSnO <sub>3</sub>      | Nanosheets                    | Na <sub>2</sub> SnO <sub>3</sub> ·3H <sub>2</sub> O/Zn(Ac) <sub>2</sub> ·2H <sub>2</sub> O/ethylene glycol                                                                                  | Solvothermal              | 71              |
| ZnSnO <sub>3</sub>      | Hollow cubes                  | ZnCl <sub>2</sub> , SnCl <sub>4</sub> ·5H <sub>2</sub> O/sodium citrate/NaOH                                                                                                                | Coprecipitation           | 116             |
| ZnSnO <sub>3</sub>      | Solid/hollow microspheres     | Zn(NO <sub>3</sub> ) <sub>2</sub> ·6H <sub>2</sub> O/SnCl <sub>4</sub> ·5H <sub>2</sub> O/NaOH                                                                                              | Precipitation             | 70              |
| ZnSnO <sub>3</sub>      | Nano-urchins                  | Zinc plate/ZnO <sub>x</sub> (OH) <sub>y</sub> /NH <sub>3</sub>                                                                                                                              | Laser ablation            | 75              |
| ZnSnO <sub>3</sub>      | Amorphous                     | ZnO/SnO <sub>2</sub>                                                                                                                                                                        | Magnetron sputtering      | 79, 141 and 142 |
| Ag-ZnSnO <sub>3</sub>   | Hollow cubes                  | ZnCl <sub>2</sub> /SnCl <sub>4</sub> ·5H <sub>2</sub> O/sodium citrate/Zn(NO <sub>3</sub> ) <sub>2</sub> ·6H <sub>2</sub> O                                                                 | Coprecipitation           | 143             |
| CDs-ZnSnO <sub>3</sub>  | Nanocubes                     | SnCl <sub>4</sub> ·5H <sub>2</sub> O/ZnSO <sub>4</sub> ·7H <sub>2</sub> O/CDs powder                                                                                                        | Precipitation-calcination | 144             |
| ZnSnO <sub>3</sub> /rGO | Nanocubes                     | ZnSO <sub>4</sub> ·7H <sub>2</sub> O/SnCl <sub>4</sub> ·5H <sub>2</sub> O/PDDA/GO                                                                                                           | Coprecipitation           | 145 and 146     |
| S-ZnSnO <sub>3</sub>    | Hollow cubes                  | Zn(NO <sub>3</sub> ) <sub>2</sub> ·6H <sub>2</sub> O, Na <sub>2</sub> SnO <sub>3</sub> ·4H <sub>2</sub> O, C <sub>4</sub> H <sub>6</sub> N <sub>2</sub> , CH <sub>3</sub> CSNH <sub>2</sub> | Ion-exchange              | 147             |
| ZnSnO <sub>3</sub> /rGO | Nanosheets                    | Zn(NO <sub>3</sub> ) <sub>2</sub> ·6H <sub>2</sub> O/SnCl <sub>2</sub> ·2H <sub>2</sub> O/GO                                                                                                | Thermal decomposition     | 76              |
| ZnSnO <sub>3</sub> /C   | Nanofibers                    | SnCl <sub>2</sub> ·2H <sub>2</sub> O/ZnCl <sub>2</sub> /DMF/PVP                                                                                                                             | Electrospinning           | 73              |



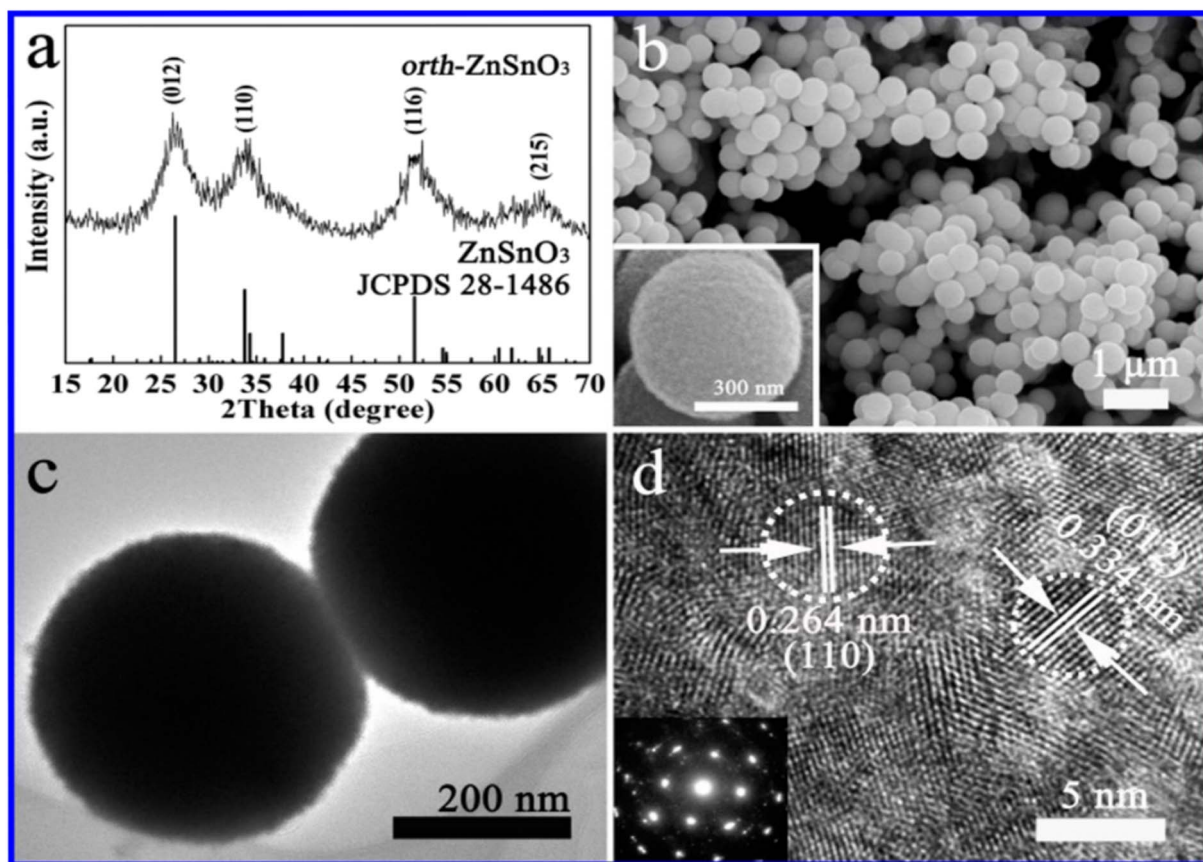


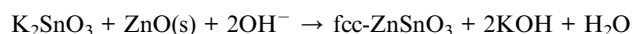
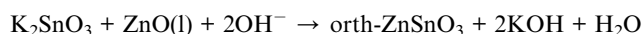
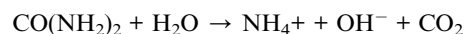
Fig. 7 (a) XRD pattern, (b) SEM image, (c) TEM image, and (d) HRTEM image of ZnSnO<sub>3</sub> solid microspheres.<sup>90</sup>

for the formation of nanoparticles in the hydrothermal synthesis process.<sup>81–85</sup>

In hydrothermal fabrication of ZnSnO<sub>3</sub>, the compositions of synthesized nanomaterials can be precisely controlled through liquid-phase or multiphase chemical reactions. An advantage of the hydrothermal method is its capability to generate crystalline phases that are not stable at higher temperatures.<sup>86–89</sup>

The hydrothermal method is commonly used for the synthesis of ZnSnO<sub>3</sub> with orthorhombic (orth) and face-centred cubic (fcc) phases. In a typical synthesis procedure, both discrete orth and fcc ZnSnO<sub>3</sub> phases are simultaneously fabricated in a single solution through hydrothermal treatment. This is achieved by treating a solution having a small amount of ZnO-covered foil, urea, and potassium stannate trihydrate under hydrothermal conditions.<sup>90</sup> According to their study, the synthesis process starts by hydrothermally treating a solution of pure zinc-foil and C<sub>20</sub>H<sub>37</sub>NaO<sub>7</sub>S, OT, and 1.50 g L<sup>−1</sup> to obtain the ZnO precursor. Subsequently, the obtained ZnO-foil is again hydrothermally treated in an alcohol–water solution containing urea (CO(NH<sub>2</sub>)<sub>2</sub>) and potassium stannate trihydrate (K<sub>2</sub>SnO<sub>3</sub>·3H<sub>2</sub>O). Subsequently, the resulting microspheres are dried in air at 60 °C for 6 h. Consequently, pure orth-ZnSnO<sub>3</sub> solid microspheres are formed in the solution, as illustrated in Fig. 7. Additionally, pure fcc-ZnSnO<sub>3</sub> hollow microspheres are observed on the ZnO-induced template, as illustrated in Fig. 8.<sup>90</sup>

The following is a description of the process for the formation of the two products:



To achieve controlled morphologies of the synthesized materials, different surfactants and polymers such as polyethylene glycol (PEG) can be introduced. Numerous studies have been published on the hydrothermal fabrication of nanoparticles, nanorods, nanotubes, nanowires, and hollow spheres. The presence of PEG in the hydrothermal reaction has been observed to facilitate the formation of nanowire morphologies.

In ref. 91, ZnSnO<sub>3</sub> nanowire arrays were synthesized using a single-step hydrothermal method. In their study, Zn(Ac)<sub>2</sub>·2H<sub>2</sub>O, SnCl<sub>4</sub>·5H<sub>2</sub>O, and polyethylene glycol (PEG) were used as the precursors. PEG, acting as a complex agent, reacted with Zn(CH<sub>3</sub>CO<sub>2</sub>)<sub>2</sub>·2H<sub>2</sub>O and SnCl<sub>4</sub>·5H<sub>2</sub>O to form a soluble ZnSn(OH)<sub>n</sub>(PEG)<sub>6−n</sub> complex. The presence of PEG was crucial for obtaining the desired morphologies of ZnSnO<sub>3</sub> nanowires. The molar ratio of Zn(CH<sub>3</sub>CO<sub>2</sub>)<sub>2</sub>·2H<sub>2</sub>O to SnCl<sub>4</sub>·5H<sub>2</sub>O to PEG



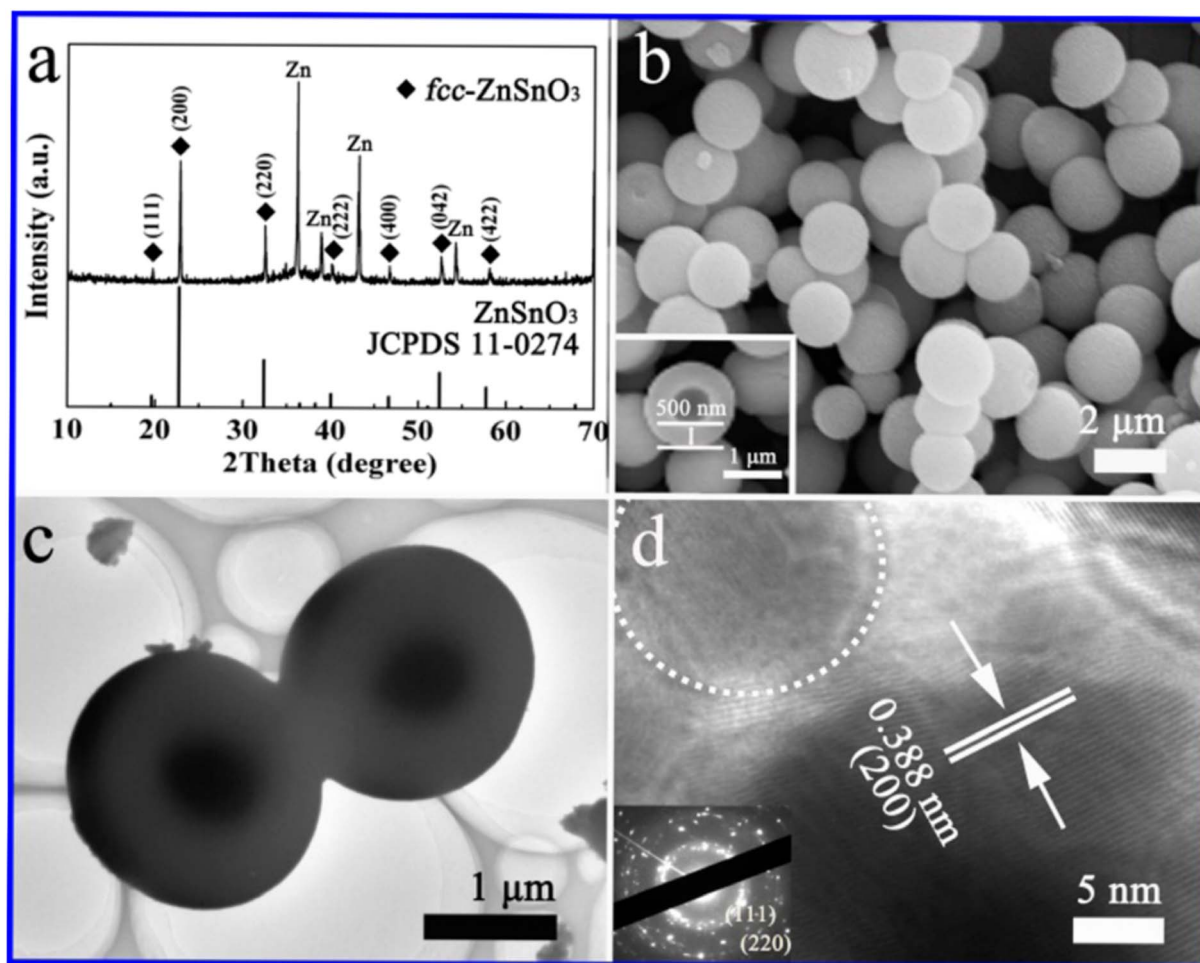
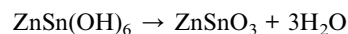


Fig. 8 (a) XRD pattern, (b) SEM image, (c) TEM image, and (d) HRTEM image of ZnSnO<sub>3</sub> hollow microspheres.<sup>90</sup>

(4000) was approximately 1:1:1, and the reaction was performed at a temperature of 200 °C for a duration of 12 h.

Hu and colleagues developed a straightforward chemical solution technique to achieve the large-scale production of well-defined faceted cubic ZnSnO<sub>3</sub> and octahedral Zn<sub>2</sub>SnO<sub>4</sub> microcrystals.<sup>92</sup> They employed an acid-etching technique, which allowed the originally synthesized zinc stannate faceted microcrystals to undergo a transformation into hollow structures, while preserving their original shape. The synthesis process involved using specific starting materials, namely, tin tetrachloride (SnCl<sub>4</sub>·5H<sub>2</sub>O), zinc acetate (ZnAc<sub>2</sub>·2H<sub>2</sub>O), and NaOH, which were diluted in distilled water to create a clear solution. In a typical procedure, a 0.02 M ZnAc<sub>2</sub> solution was added to a 0.02 M SnCl<sub>4</sub> solution at RT and vigorously agitated for 10 min to form a mixed solution for the synthesis of cubic ZnSnO<sub>3</sub> microcrystals. Subsequently, this mixture was combined with a 0.2 M NaOH solution (15 mL) and continuously stirred for an additional 10 min. The proportion of Zn<sup>2+</sup>:Sn<sup>4+</sup>:Na<sup>+</sup> was kept at a molar ratio of 1:1:10. Subsequently, a hydrothermal procedure was conducted at 130 °C for 6 h inside a reactor with a 60 mL capacity. After the reaction, the resulting white materials were washed multiple times with ethanol and distilled water.

The reactions that lead to the creation of ZnSnO<sub>3</sub> can be summarized as follows:<sup>92</sup>



Under hydrothermal conditions, the creation, and then breaking of the weak-phase ZnSn(OH)<sub>6</sub> results in the nucleation and development of ZnSnO<sub>3</sub> nanoparticles, involving the aforementioned chemical processes (step 1). More ZnSnO<sub>3</sub> nanocrystallites are generated after adding more reactants to the reaction mixture, which are further combined and bound into comparatively augmented nanocrystallites. Noticeably, the obtained ZnSnO<sub>3</sub> undergoes a dissolution-recrystallisation process to reduce the high-energy surfaces. Next, the particles combine to form larger aggregates, adhering to the principles of Ostwald ripening (step 2). By extending the reaction duration, the crystals undergo morphological changes and acquire a cubic shape with dimensions measuring several microns. The cubic crystals exhibit 100 lattice planes on their basal surfaces,



# Growth Process

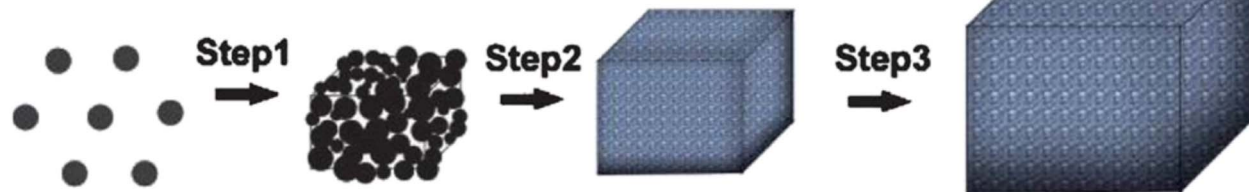


Fig. 9 Possible mechanism for the growth of  $\text{ZnSnO}_3$  cubic microcrystals.<sup>92</sup>

achieved by adjusting the growth conditions to promote the desired crystal facet (step 3). Lastly, the nanocrystals present in the reaction system adhere to the surfaces of the  $\text{ZnSnO}_3$  cubic crystals, resulting in the production of huge cubes.<sup>92</sup> Fig. 9 schematically depicts the growth process.

Exercising meticulous control over the chemical characteristics of the mineralizer is crucial, given that it determines the specific surface charges present on the resulting metal oxides. These surface charges serve as a critical factor in the overall process, significantly influencing and facilitating the creation of nanoparticles during the hydrothermal synthesis procedure. Generally, the careful management of the properties of the mineralizer directly impacts the outcome of nanoparticle formation, underscoring its pivotal role in this intricate process. However, the reliability and reproducibility of the process are limited, the necessary equipment is expensive, a longer reaction time is required, and it consumes plenty energy.

## 3.2. Solvothermal

The solvothermal procedure is similar to the hydrothermal method, excepted for the use of organic solvents instead of

water. In the case of alcohols and glycerol as the reaction medium, the reactions are referred to as alcohothermal and gyrothermal, respectively. For the synthesis of NCs with good crystalline characteristics, these synthetic techniques are crucial.<sup>93</sup>

Wang and colleagues used the solvothermal route to achieve a phase transformation from 3D fcc  $\text{ZnSnO}_3$  nanosheets to 2D orth  $\text{ZnSnO}_3$  nanosheets. In the conventional route, a white suspension of  $\text{ZnSn(OH)}_6$  products was promptly formed by combining a solution of  $\text{Na}_2\text{SnO}_3 \cdot 3\text{H}_2\text{O}$  and  $\text{Zn}(\text{Ac})_2 \cdot \text{H}_2\text{O}$  in a mixture of ethylene glycol–deionized water. Subsequently, the  $\text{ZnSn(OH)}_6$  was moved to a stainless-steel Teflon-lined autoclave and kept at 180 °C for 12 h. Next, the obtain powder was annealed at 500 °C for 4 h at a rate of 2.8 °C  $\text{min}^{-1}$ . Finally, two-dimensional  $\text{ZnSnO}_3$  nanosheets were created.<sup>71</sup>

With the shift from 3D nanocubes to 2D  $\text{ZnSnO}_3$  nanosheets, a straightforward solvothermal pathway was investigated in ref. 94. To form aqueous solutions,  $\text{Na}_2\text{SnO}_3 \cdot 3\text{H}_2\text{O}$  and  $\text{Zn}(\text{Ac})_2 \cdot 3\text{H}_2\text{O}$  were mixed in ethanol–water solution in the usual manner. Subsequently, the  $\text{Na}_2\text{SnO}_3$  solution was added slowly to the  $\text{Zn}(\text{CH}_3\text{COO})_2$  solution with stirring, resulting in the

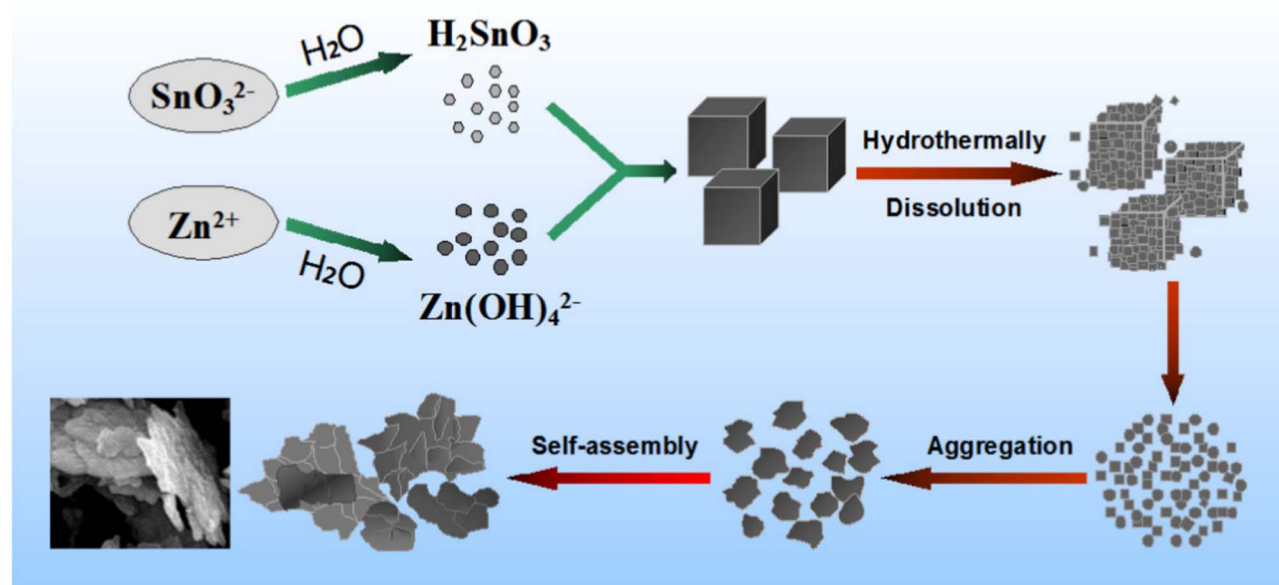
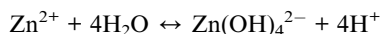
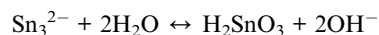


Fig. 10 Diagram depicting the probable process for the formation of the precursor  $\text{ZnSn(OH)}_6$ .<sup>95</sup>



formation of a white  $\text{ZnSn}(\text{OH})_6$  suspension. Then, the suspension was transferred to a Teflon-lined autoclave and kept at 180 °C for 12 h. The resulting  $\text{ZnSn}(\text{OH})_6$  was washed and annealed at 600 °C for 3 h. Firstly, the reaction happened due to the hydrolysis of the  $\text{SnO}_3^{2-}$  ions, resulting in the creation of blend  $\text{H}_2\text{SnO}_3$ . Later, the  $\text{Zn}^{2+}$  ions undergo a reaction with OH ions, resulting in the formation of the  $\text{Zn}(\text{OH})_4^{2-}$  phase. This reaction occurs due to the hydrolysis of both  $\text{CH}_3\text{COO}^-$  and  $\text{Zn}^{2+}$  ions, as follows:<sup>94</sup>



Under hydrothermal conditions, the metastable  $\text{ZnSn}(\text{OH})_6$  nanocubes will undergo decomposition and subsequent recrystallization as the temperature and pressure increases. This process follows the breaking-recrystallization pathway, resulting in the formation of a stable nanostructure. Fig. 10 depicts the hypothetical growth process graphically.

### 3.3. Sol-gel technique

The sol-gel method is a chemical process that creates oxide-based materials from hydrolysable precursors through hydrolysis and condensation reactions. These precursors contain weaker ligands than water, such as halides, nitrates, sulfates,

alkoxides, and carboxylates. The hydrolysed precursors form small colloidal nanoparticles in a liquid sol, which can undergo further polycondensation to create a network of polymeric oxide-based materials with oxo-bridges. The initial gels formed through this method consist of both a gel network and a significant liquid phase. Drying these gels, whether at room temperature or through heating, removes the solvent phase and produces dense materials.<sup>96–98</sup>

Due to its simplicity, cost-effectiveness, and capability to produce large-area films, the sol-gel method is widely regarded as an excellent technique for the preparation of  $\text{ZnSnO}_3$ . As demonstrated in ref. 99, the face-centred trigonal perovskite structure of  $\text{ZnSnO}_3$  was synthesized using a sol gel. They used  $\text{Na}_2\text{SnO}_3 \cdot 3\text{H}_2\text{O}$  (0.01 mol) and  $\text{Zn}(\text{NO}_3)_2 \cdot 6\text{H}_2\text{O}$  (0.01 mol) as the precursors and deionized water (100 mL) as the solvent. Finally, the white  $\text{ZnSnO}_3$  nanoparticles were dried at 70 °C in an oven for 12 h.

The authors of ref. 100 fabricated polycrystalline IL-type  $\text{ZnSnO}_3$  via the sol-gel process at atmospheric pressure utilizing  $\text{Zn}(\text{CH}_3\text{COO})_2 \cdot 2\text{H}_2\text{O}$ ,  $\text{SnCl}_4 \cdot 5\text{H}_2\text{O}$ , and ethylene glycol as the precursors. To determine the impact of sintering on the  $\text{ZnSnO}_3$  structure, the resulting product was sintered at 450 °C, 550 °C, and 650 °C. Upon sintering below 500 °C, they observed that the phase formation is incomplete with the significant presence of ZnO and  $\text{SnO}_2$  as secondary phases. In contrast, sintering at 650 °C led to the creation of pure-phase  $\text{ZnSnO}_3$ .

Li *et al.* synthesized Zn–Sn–O thin-films and investigated the impact of sintering (300–1000 °C) on their microstructure, morphological, and optical characteristics.<sup>67</sup> Spinel  $\text{Zn}_2\text{SnO}_4$

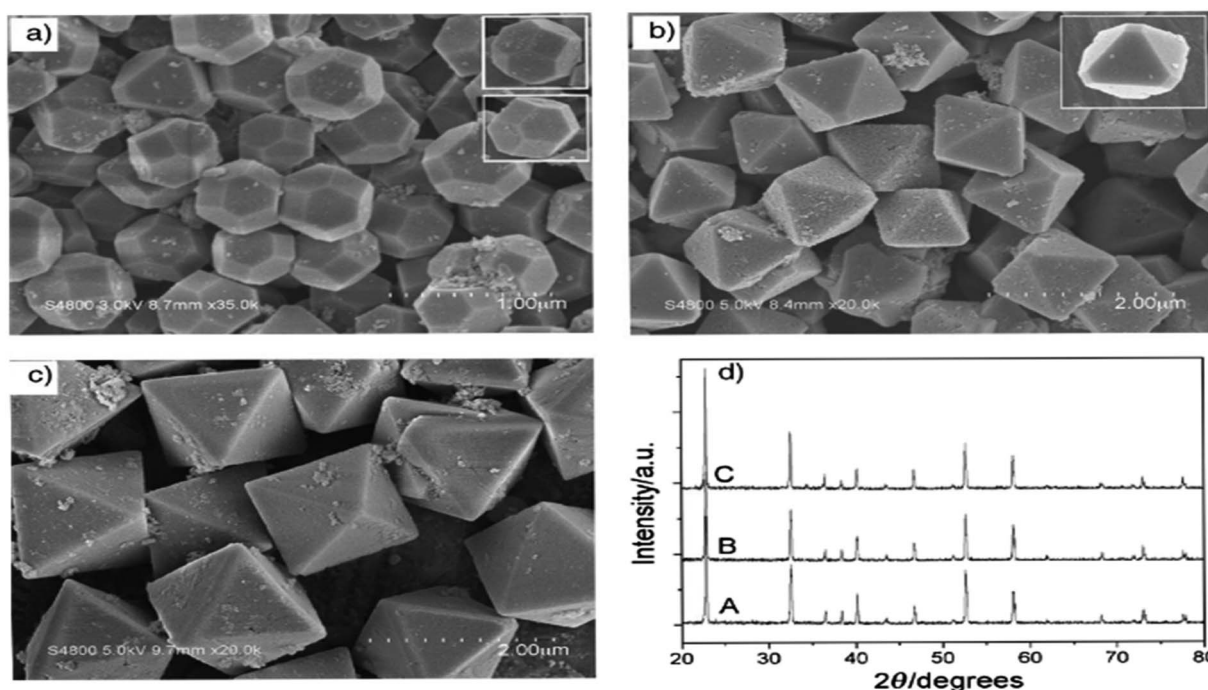
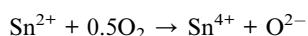


Fig. 11 SEM micrograph of  $\text{ZnSnO}_3$  obtained using: (a) SDBS 0.15 M, (b) SDBS 0.4 M, and (c) SDBS 0.75 M. (d) XRD pattern of (A) 14-faceted polyhedra, (B) truncated octahedra and (C) octahedra.<sup>102</sup>

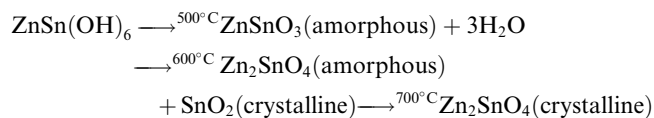


was formed in the temperature range of 400–700 °C, while perovskite  $\text{ZnSnO}_3$  emerged at 800 °C. Alternatively, the formation of  $\text{Zn}_2\text{SnO}_4$  only occurred at temperatures exceeding 1000 °C. Furthermore, the grain size decreased as the temperature increased from 400 °C to 800 °C. The proposed explanation is that the  $\text{Zn}_2\text{SnO}_4$  grain grows *via* the surface diffusion pathway, and  $\text{SnO}_2$  is dispersed on the surface of  $\text{Zn}_2\text{SnO}_4$ , inhibiting grain formation. Furthermore, when the temperature exceeds 800 °C, the lattice properties clearly decrease. The formation of  $\text{ZnSnO}_3$  is responsible for this phenomenon, as follows:



Authors of ref. 101 synthesized porous  $\text{ZnSnO}_3$  nanocubes *via* the conventional solution-based process together with a calcination process. When annealed at 500 °C,  $\text{ZnSn(OH)}_6$  lost three water molecules to form amorphous  $\text{ZnSnO}_3$ , which decomposed into amorphous  $\text{Zn}_2\text{SnO}_4$  and crystalline  $\text{SnO}_2$  at

600 °C. The amorphous  $\text{Zn}_2\text{SnO}_4$  became crystalline  $\text{Zn}_2\text{SnO}_4$  when the annealing temperature increased to 700 °C. The proposed conversion of  $\text{ZnSn(OH)}_6$  into different morphologies is as follows:



As demonstrated in ref. 103,  $\text{ZnSnO}_3$  nanostructures were prepared *via* the sol-gel method using TEPA as the gelling medium and trimetric acid hydrolysis medium, which increased the hydrolysis rate and led to initially smaller nucleation, controlling the final particle size. Here, 0.8 mmol  $\text{ZnCl}_2$  and 0.8 mmol  $\text{SnCl}_2$  were dispersed in 10 mL DI individually, and then mixed with the gelling agent and hydrolysis agent at 80 °C. Subsequently, the fabricated  $\text{ZnSnO}_3$  was calcined at 700 °C for 2 h. A straightforward and cost-effective method was employed by Geng *et al.* for the successful synthesis of polyhedral  $\text{ZnSnO}_3$  with a variety of morphologies without further heat treatment.<sup>32</sup> They showed the

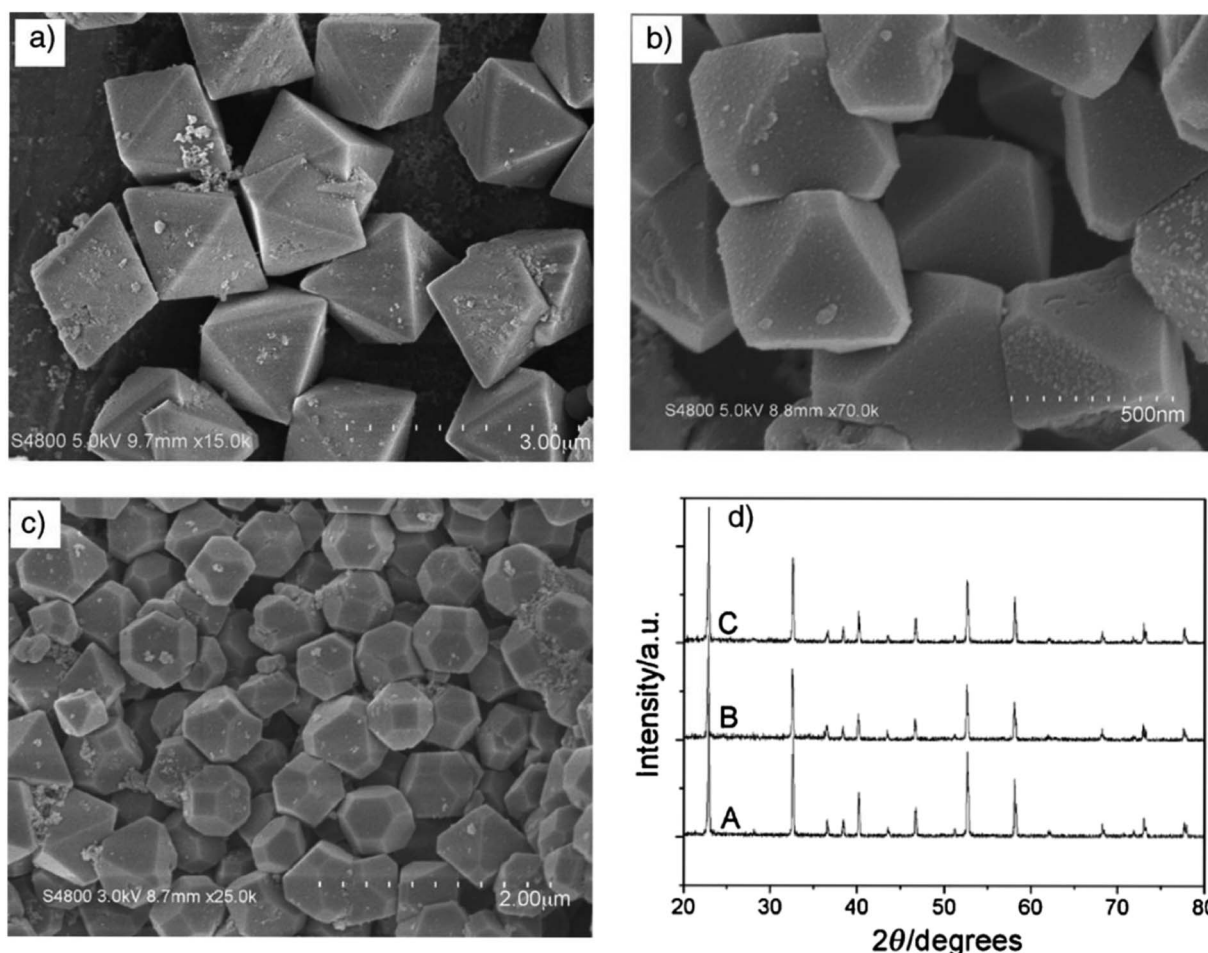


Fig. 12 Typical SEM images of the as-prepared  $\text{ZnSnO}_3$  products: (a) CTAB 0.15 M, (b) CTAB 0.4 M, and (c) CTAB 0.75 M. (d) Corresponding XRD patterns of the as-prepared  $\text{ZnSnO}_3$  polyhedra, (A) octahedra, (B) truncated octahedra and (C) 14-faceted polyhedra.<sup>102</sup>

morphology conversion between octahedra and 14-faceted  $\text{ZnSnO}_3$  using surfactant. In the conventional synthesis, they used tin tetrachloride and  $\text{Zn}(\text{Ac})_2$  as the starting precursors. Different surfactants in varying concentrations and NaOH were used to control the polyhedral shape. In a beaker, the combined sol was treated at 85 °C with stirring. Finally, centrifugation was used to collect the white polyhedral  $\text{ZnSnO}_3$  products. When an anionic surfactant (SDBS) was employed, the synthesis of  $\text{ZnSnO}_3$  yielded different polyhedral shapes depending on the concentration of SDBS such as (1) at 0.25 M, 14-faceted polyhedra were synthesized, (2) at a moderate range of 0.3 to 0.5 M, truncated octahedra were formed, and (3) at 0.5 to 1.0 M, typical octahedra were synthesized. These observations demonstrate how the SDBS concentration influences the morphology of the resulting  $\text{ZnSnO}_3$  particles. Fig. 11 depicts this shape-evolution pattern. Again, at varying concentrations of the cationic CTAB surfactant, the synthesis of  $\text{ZnSnO}_3$  yielded different polyhedral shapes such as (1) at 0.2 M, regular octahedra were created, (2) at an intermediate CTAB range of 0.3 to 0.5 M, truncated octahedra were formed, and (3) at 0.5 to 1.0 M, the resulting products exhibited regular 14-faceted polyhedra. These findings demonstrate the influence of CTAB concentration on the morphology of the synthesized  $\text{ZnSnO}_3$  particles. This shape-evolution form is depicted in Fig. 12.

When the concentration of the anionic SDBS surfactant increased, the  $R$  value also increased. Consequently, the shape of the  $\text{ZnSnO}_3$  crystals changed from 14-faceted polyhedra to truncated octahedra, and eventually to regular octahedra. In

contrast, when the concentration of the cationic CTAB surfactant increased, the  $R$  value decreased. This resulted in a different trend compared to the anionic SDBS surfactant, where increasing the CTAB concentration led to a transition from regular octahedra to truncated octahedra, and ultimately to 14-faceted polyhedra (Fig. 13).<sup>102</sup>

The production of  $\text{ZnSnO}_3$  using the sol-gel method presents significant technical and chemical challenges. This is primarily due to the fact that the precursors used to introduce  $\text{Sn}^{4+}$  ions typically contain highly electronegative ions (such as  $\text{Cl}^-$ ) or additional metallic ions (such as  $\text{Na}^{2+}$ ), which reduce the likelihood of effective bonding between  $\text{Sn}^{4+}$  and  $\text{Zn}^{2+}$ . Additionally,  $\text{ZnSnO}_3$  exhibits lower thermal stability, leading to its decomposition into  $\text{Zn}_2\text{SnO}_4$ ,  $\text{ZnO}$ , and  $\text{SnO}_2$  at elevated temperatures. Thus, to address these issues, alternative low-temperature synthesis techniques such as co-precipitation and hydrothermal synthesis have been employed.<sup>104</sup>

### 3.4. Vapor-liquid-solid technique

In recent decades, the vapor-liquid-solid (VLS) growth technique has witnessed substantial advancements.<sup>105</sup> Originally used for whisker growth, it has evolved into a practical method for producing semiconductor nanowires. These nanowires have applications in various fields such as nanoscale electronics, optoelectronics, sensing, and energy conversion.<sup>106–108</sup> However, based on a previous literature survey, few research endeavours have been reported on the synthesis of  $\text{ZnSnO}_3$  using the VLS method.

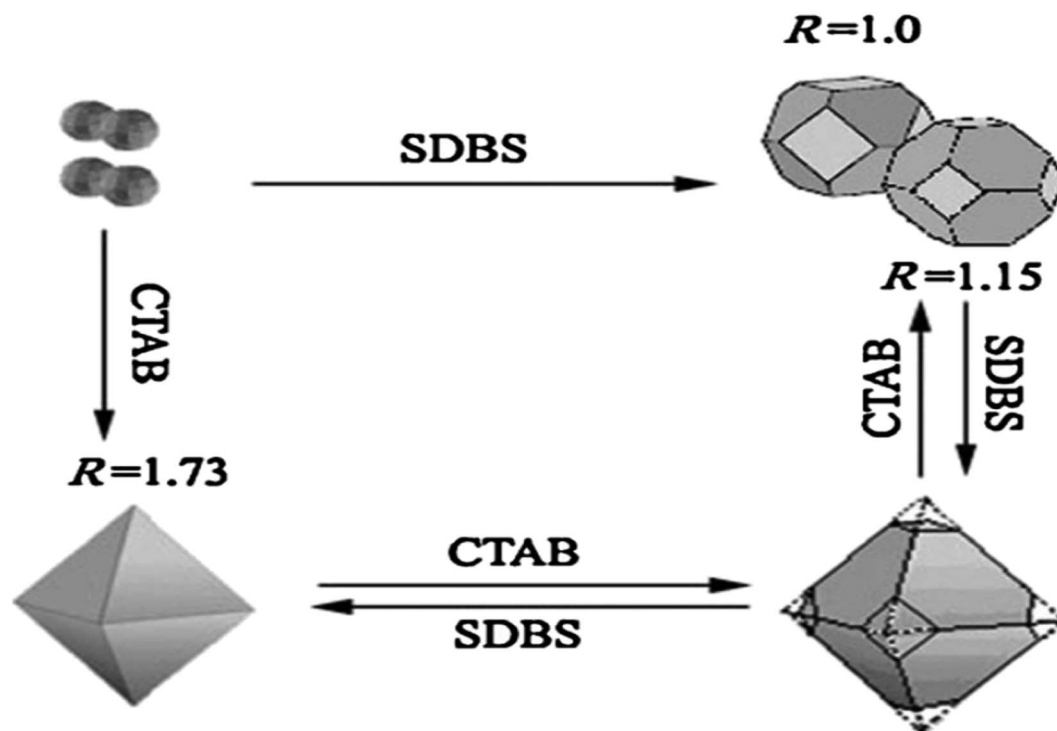


Fig. 13 Schematic diagram depicting the relationship between the ratio  $R$  and the transformation of the crystal shape.<sup>102</sup>



Wang *et al.* employed the widely recognized vapor-liquid-solid (VLS) mechanism to establish a thermal reaction method for synthesizing  $\text{ZnSnO}_3$  nanoparticles, as described in ref. 36. The typical synthesis procedure utilized in their study is as follows. Initially, a Zn-Sn buffer layer was deposited as a seed layer on Si and alumina substrates. Subsequently, separate starting materials consisting of Zn and Sn powders were set in an alumina boat. Moreover, to enhance the carbon-thermal reaction while sintering, a blend of source materials containing 50 wt% graphite powder was placed in the central region of a quartz reactor with a stable temperature of 1173 K and left to react at a pressure of 10–20 torr for a duration of 3 h. Throughout the reaction, the flow rate of argon and oxygen gases was maintained at 200 sccm and 10–20 sccm, respectively. Subsequently, the resulting product comprised of single microbelt nanogenerators was constructed using enlarged  $\text{ZnSnO}_3$  belts measuring 300–1000  $\mu\text{m}$  in length. Furthermore, the power output and piezoelectric properties of these microbelts were assessed.<sup>36</sup>

Wang *et al.* also synthesized lead-free  $\text{ZnSnO}_3$  triangular-belts using a high-temperature carbon-thermal approach. In this method, the starting materials consisted of Zn and Sn together with graphite powder. These materials were set in the middle point of the reactor to facilitate the process. The system was left to react at a stable temperature of 900 °C and a pressure of 10–20 torr for a duration of 3 h. The flow rate of Ar and  $\text{O}_2$  gas during the reaction process was set at 200 and 10–20 sccm, respectively.<sup>37</sup>

Although the vapor-liquid-solid (VLS) technique is a potent method for the synthesis of nanowires, it is associated with certain challenges. For example, it requires specialized, costly equipment for handling the potentially hazardous precursor gases, and achieving uniformity in nanowire properties can be difficult. Furthermore, the incorporation of contaminants during growth process and the energy-intensive high-temperature conditions are additional concerns,<sup>109–111</sup> together with its limited scalability and precursor availability. Nevertheless, VLS is valuable due to its precise control of the

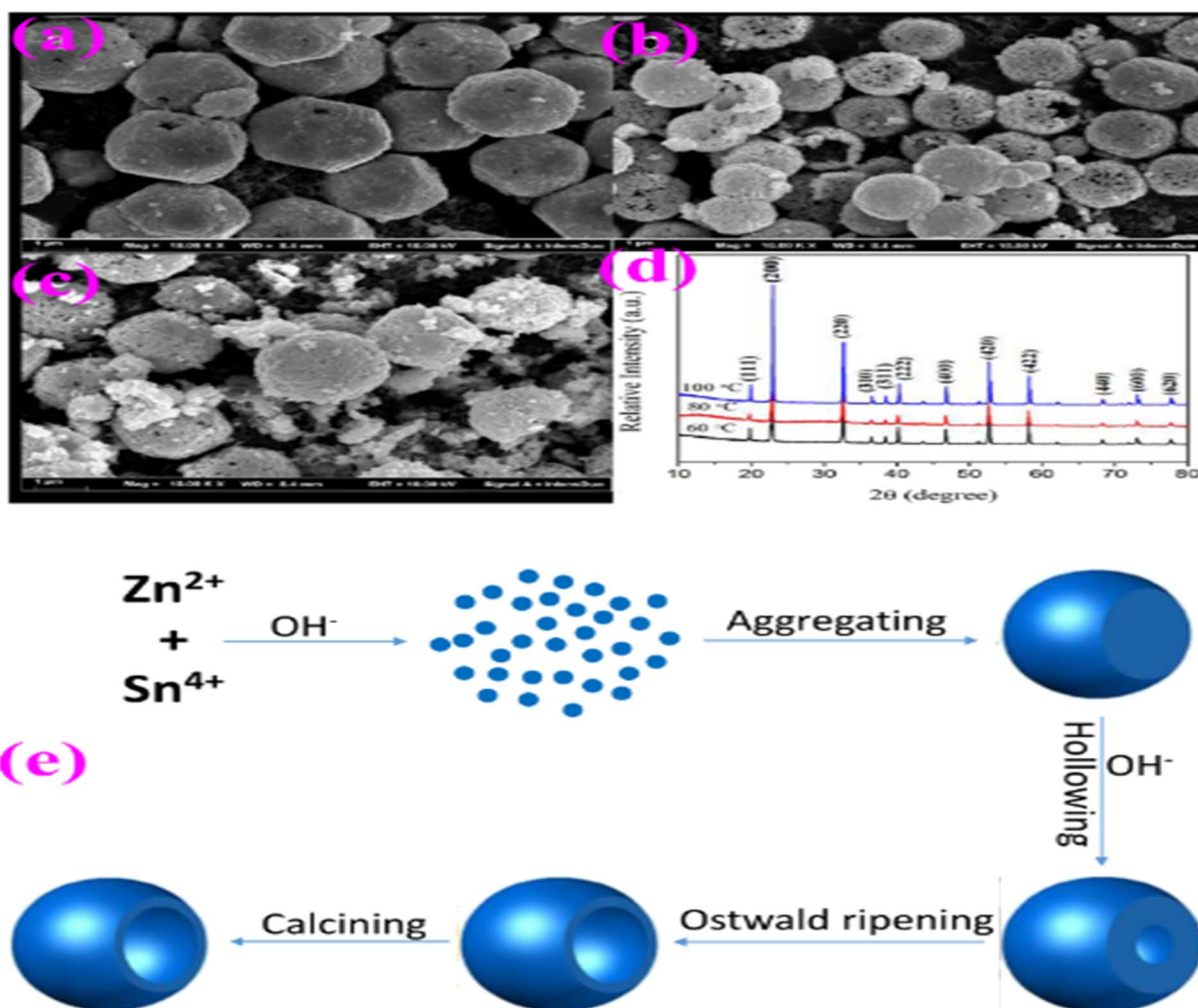


Fig. 14 SEM images of  $\text{ZnSnO}_3$  obtained at different reaction temperatures: (a) 60 °C, (b) 80 °C, and (c) 100 °C. (d) XRD patterns of  $\text{ZnSn(OH)}_6$  precursor. (e) Schematic illustration of the possible formation mechanism of  $\text{ZnSnO}_3$  hollow spheres.<sup>70</sup>

nanowire properties, and thus ongoing research seeks to address the above-mentioned challenges for its broader utilization.

### 3.5. Coprecipitation technique

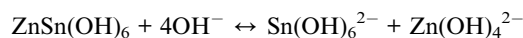
The coprecipitation method offers several advantages for the synthesis of  $\text{ZnSnO}_3$ . Firstly, it is a relatively simple and cost-effective technique, which facilitates precise control of the composition and stoichiometry of the final product. This method enables the simultaneous precipitation of zinc and tin ions in a homogeneous manner, leading to a well-mixed precursor solution.<sup>112–114</sup> Additionally, coprecipitation is conducted in aqueous solutions at relatively low temperatures, reducing the energy consumption and environmental impact compared to high-temperature synthesis methods.<sup>115</sup> The coprecipitation technique involves the precipitation of metal hydroxides from a salt precursor using a base in a solvent. By carefully controlling the release of anions and cations, the nucleation and growth kinetics of the particles can be regulated, leading to the synthesis of uniformly dispersed nanoparticles. Proper regulation of the experimental parameters such as pH, reactant and ion concentration, and temperature is crucial for controlling the morphological characteristics during the precipitation process.

In a study conducted by the authors of ref. 116, double-shelled perovskite  $\text{ZnSn(OH)}_6$  hollow-cubes were fabricated using a rapid coprecipitation method. This involved mixing  $\text{Zn}^{2+}$ ,  $\text{Sn}^{4+}$ , and  $\text{OH}^-$  ions with  $\text{Na}_3\text{C}_6\text{H}_5\text{O}_7$  and  $\text{NaOH}$  at RT. Subsequently, the obtained  $\text{ZnSn(OH)}_6$  cubes were transformed into  $\text{ZnSnO}_3$  double-shelled nanocubes through sintering at 300 °C for 3 h in an  $\text{N}_2$  atmosphere.

According to ref. 70,  $\text{ZnSnO}_3$  hollow spheres were fabricated via a straightforward and template-free *in situ* precipitation technique, followed by a dehydration step. During this process, a solution comprised of a water–alcohol mixture was created and subjected to constant stirring. The solution consisted of  $\text{Zn(NO}_3)_2 \cdot 6\text{H}_2\text{O}$ ,  $\text{SnCl}_4 \cdot 5\text{H}_2\text{O}$ , and  $\text{NaOH}$  mixed with the above-mentioned solution. The experimental conditions were kept constant throughout, except for the growth temperature, which was varied from 60 °C to 100 °C to investigate its impact on the morphology of  $\text{ZnSnO}_3$ . The results showed a significant alteration in the shape of  $\text{ZnSnO}_3$  with an increase in the growth temperature, as depicted in Fig. 14. At 60 °C, the sample primarily consisted of uniform and monodisperse polyhedrons with a diameter of approximately 2.4  $\mu\text{m}$  (Fig. 12a). When the growth temperature was increased to 80 °C, the  $\text{ZnSnO}_3$  nanoparticles assembled into hollow microspheres (Fig. 14b). However, at 100 °C, uneven  $\text{ZnSnO}_3$  spheres were obtained, with some small nanoparticles attached to their surface (Fig. 14c). Fig. 14d displays the X-ray diffraction peaks of the  $\text{ZnSn(OH)}_6$  precursor at various growth temperatures.

A schematic representation of the fabrication route of the  $\text{ZnSnO}_3$  hollow-spheres is shown in Fig. 14e. Initially, a significant amount of  $\text{ZnSn(OH)}_6$  is generated in the presence of  $\text{Zn}^{2+}$  and  $\text{Sn}^{4+}$  ions at high concentrations of  $\text{OH}^-$ . Subsequently, these  $\text{ZnSn(OH)}_6$  particles develop into spherical shapes

through rapid aggregation. In the second stage, the surface of the  $\text{ZnSn(OH)}_6$  microspheres undergo etching due to the dissolution approach in the highly alkaline environment, as the concentration of  $\text{OH}^-$  ions increases. Concurrently, the re-crystallization of  $\text{ZnSn(OH)}_6$  occurs in the system, maintaining dynamic-equilibrium.<sup>70</sup>



The resulting  $\text{ZnSnO}_3$  nanoparticles typically exhibit high purity, fine particle size, and enhanced crystallinity, making coprecipitation an attractive choice for the production of  $\text{ZnSnO}_3$  for various applications, including catalysis, sensors, and photovoltaics.<sup>112,115</sup> However, one drawback of the coprecipitation method for synthesizing  $\text{ZnSnO}_3$  is controlling the particle size and morphology of the resulting material. This method often yields nanoparticles with a wide size distribution and irregular shapes, which can negatively impact the properties and performance of the materials in specific applications.<sup>117,118</sup> Achieving precise control over the particle size and shape can be difficult, requiring additional steps such as post-synthesis annealing and the use of surfactants.<sup>119</sup> Moreover, coprecipitation may also introduce impurities or defects in the final product, which can be undesirable in some applications, such as electronic devices and photocatalysis, where material purity and uniformity are critical for optimal performance.

### 3.6. Magnetron sputtering technique

Sputtering involves bombarding a target material with high-energy particles, causing the discharge of atoms/molecules from the surface. However, diode sputtering has notable drawbacks, including low deposition rates and high cost. In contrast, magnetron sputtering is a vacuum-coating process known for its high deposition rates, allowing the deposition of metals, alloys, and compounds on various substrates with thicknesses up to the millimeter scale. This technique offers several significant advantages compared to other vacuum coating processes, making it suitable for diverse commercial applications ranging from microelectronic manufacturing to the creation of decorative coatings.<sup>120</sup>

$\text{ZnSnO}_3$  films were produced using variable magnetron sputtering settings, employing mixed powder targets of  $\text{ZnO}$  and  $\text{SnO}_2$ . In the resulting films,  $\text{ZnSnO}_3$  nanocrystalline phases were preferentially formed in the columnar grain structure when a short deposition time of 2 h was used. To create the target, pure  $\text{SnO}_2$  and  $\text{ZnO}$  particles with a size of 1  $\mu\text{m}$  were blended in a rotatable drum, maintaining an atomic ratio of  $\text{Zn}/\text{Sn} = 1$ . Subsequently, the mixture was placed on a copper plate, ensuring uniform thickness and a compacted surface. The deposition process involved evacuating the system to a base pressure below  $3 \times 10^{-3}$  Pa, followed by backfilling with Ar gas to achieve a pressure of 0.1–0.5 Pa, depending on the specific array parameters. The glass made of sodium as the substrate was subjected to an in-place cleaning process through RF sputtering for 15 min at 100 W prior to deposition.



Subsequently, the films were deposited *via* RF sputtering for 4 h, using target powers in the range of 200 to 400 W and substrate spacing in the range of 80 to 180 mm based on the specific line-up run surroundings. The deposition rate of the film thickness increased with a variation in the power and spacing, obtaining improved optical properties. The average transmittance of the ZnSnO<sub>3</sub> films was about 80%. Moreover, the ZnSnO<sub>3</sub> films exhibited optical band gaps in the range of 2.6–3.4 eV and resistivity in the range of 10<sup>−3</sup>–10<sup>−4</sup> Ω cm.<sup>79</sup>

A drawback associated with the magnetron sputtering technique for the synthesis of ZnSnO<sub>3</sub> is its restricted capacity to accurately control the stoichiometry and composition. This method relies on the sputtering of target materials (in this case, Zn and Sn) on a substrate, and achieving the desired stoichiometry can be challenging due to the differences in the sputter rates between the elements. Consequently, achieving a precise ZnSnO<sub>3</sub> composition may require careful process optimization and monitoring, which can be time-consuming and may lead to variations in material quality. Additionally, magnetron sputtering is typically conducted in a vacuum environment, which can be expensive and may limit the scalability of the production process for large-scale applications.

### 3.7. Electrospinning technique

Standard single-needle electrospinning processes have demonstrated significant performance advantages and wide-ranging applications in various sectors, due to their unique structural characteristics. Electrospinning is influenced by factors such as viscosity, operating voltage, temperature, pressure, and flow velocity.<sup>121</sup>

To prepare the electrospinning precursor solution, tin(II) chloride dihydrate (SnCl<sub>2</sub>·H<sub>2</sub>O) and zinc chloride (ZnCl<sub>2</sub>) were dispersed in a mixture of dimethylformamide and absolute ethanol with vigorous stirring, maintaining a Zn<sup>2+</sup> to Sn<sup>4+</sup> molar ratio of 1:1. Subsequently, polyvinylpyrrolidone (PVP) was added to the mixture, while continuously stirring. Then, the resulting solution was transferred to a plastic syringe with a volume of 10 mL, which was outfitted with a blunt-tip needle of 22-gauge size, serving as the electrospinning precursor solution. A syringe pump was employed to control the feeding

rate. Aluminium foil was positioned 20 cm away from the needle to collect the electrospun fibers. To fabricate ZnSnO<sub>3</sub> nanofibers, the collected precursor fibers were subjected to calcination in an air environment at 450 °C for a duration of 24 h.<sup>73</sup>

An inherent limitation when employing the electrospinning technique for the synthesis of ZnSnO<sub>3</sub> is the potential difficulty in achieving the desired crystalline structure and phase purity. Electrospinning primarily produces materials in the form of nanofibers or nanofibrous mats, which may require additional heat treatment or annealing steps to transform them into the desired ZnSnO<sub>3</sub> crystalline structure. This post-treatment can introduce challenges in terms of controlling the grain size and phase purity of the final product, given that the annealing conditions need to be carefully optimized.

### 3.8. Etching technique

An advantage of the etching technique for the synthesis of ZnSnO<sub>3</sub> is its ability to create finely tailored nanostructures with a high degree of precision.<sup>122</sup> By selectively removing certain components from the precursor material, etching enables the fabrication of intricate and well-defined nanoarchitectures of ZnSnO<sub>3</sub>, such as nanofibers.<sup>73</sup> Furthermore, the etching technique enables the incorporation of ZnSnO<sub>3</sub> into specific device architectures, offering versatility and customization in the design of advanced materials for various technological applications.<sup>73,122,123</sup>

A successful etching approach was employed to produce hollow ZnSnO<sub>3</sub> nanocube architectures. Depending on the type of etching agent utilized, two paths can be distinguished, *i.e.*, acid and basic agents. In the case of utilizing NaOH solution as both the reactant and etching agent, hollow ZnSnO<sub>3</sub> architectures can be synthesized.<sup>124</sup> Initially, solid ZnSn(OH)<sub>6</sub> cubic compounds were formed by combining SnCl<sub>4</sub>/ZnCl<sub>2</sub>. Subsequently, ZnSn(OH)<sub>6</sub> hollow cube structures were obtained by introducing additional NaOH. Finally, the ZnSn(OH)<sub>6</sub> hollow cubes were sintered at an elevated temperature to form hollow ZnSnO<sub>3</sub>. Fig. 15 illustrates the progression of the hollow ZnSnO<sub>3</sub> microstructure. In further investigations, Xu *et al.* discovered that only the combination

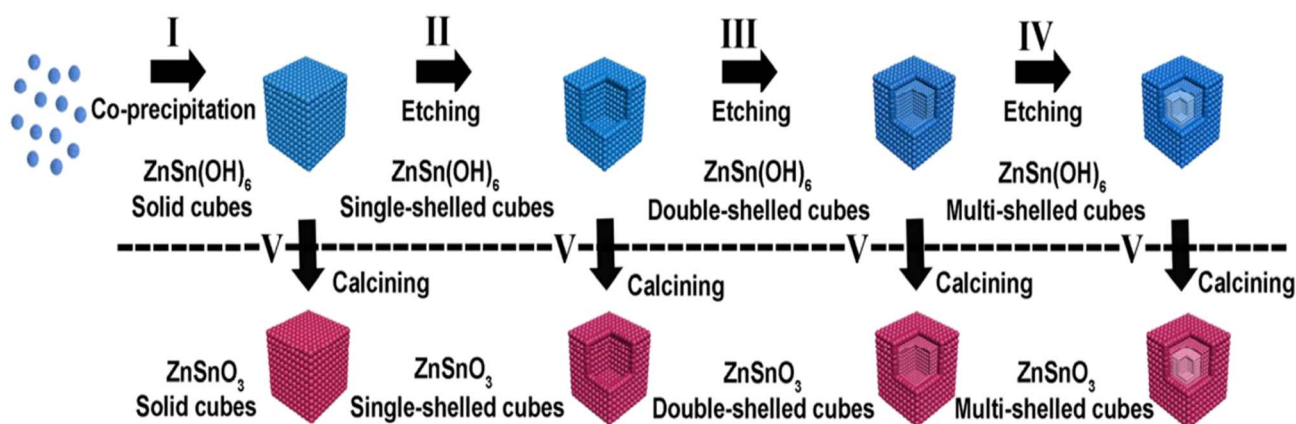


Fig. 15 Schematic illustration of the fabrication of ZnSnO<sub>3</sub> hollow-boxes under basic-etching and calcination.<sup>129</sup>

## Etching Process

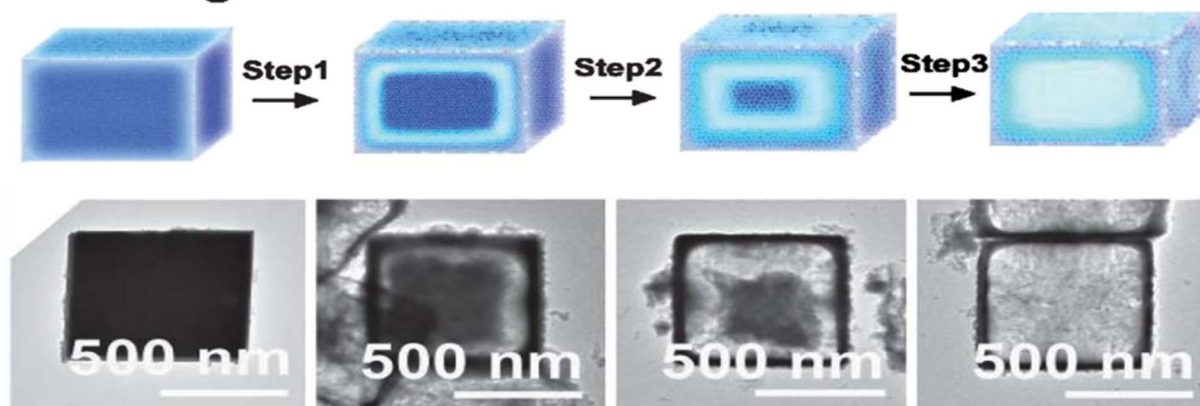


Fig. 16 Potential mechanism illustrating the acid etching process and the resulting synthesized products when employing cubic  $\text{ZnSnO}_3$  templates.<sup>92</sup>

of  $\text{SnCl}_4/\text{ZnCl}_2/\text{NaOH}$  solution allowed the fabrication of  $\text{ZnSnO}_3$  hollow-boxes.<sup>125</sup> Similarly, by subjecting a mixture of  $\text{SnCl}_4/\text{Zn}(\text{Ac})_2/\text{NaOH}$  to hydrothermal heating based on Ostwald ripening, hierarchically hollow  $\text{ZnSnO}_3$  cages were synthesized in a one-pot reaction, requiring an excess amount of  $\text{NaOH}$ .<sup>126–128</sup>

Cubic microcrystals of  $\text{ZnSnO}_3$  can be readily transformed into hollow structures by utilizing nitric acid ( $\text{HNO}_3$ ) as an alternative etching agent. Fig. 16 presents the proposed mechanism of the acid etching process, together with TEM images depicting the products obtained at different stages of etching the cubic  $\text{ZnSnO}_3$  templates.<sup>92</sup>

The etching technique for the synthesis of  $\text{ZnSnO}_3$  has drawbacks, including challenges in controlling the composition and morphology precisely. Etching involves selectively removing components from a precursor material, which is difficult for complex oxide systems such as  $\text{ZnSnO}_3$ . This can lead to non-uniformity in the final product, affecting the properties and performance. The process is time-consuming, requiring careful optimization of the etching agents and conditions, making synthesis more intricate compared to other methods.

### 3.9. Mechanochemical technique

Mechanochemical blending is a solid preparation method involving the coupling of mechanical and chemical phenomena on the molecular scale. It is an efficient strategy for synthesizing nanosized metal lattice composite powders due to its simplicity and the ability to achieve composite powder particles with a uniform distribution of grain sizes. In the study reported in ref. 130, the orthorhombic structure of  $\text{ZnSnO}_3$  was created through mechanochemical grinding. A stoichiometric amount of  $\text{ZnO}$  and  $\text{SnO}_2$  was ground for 20 min to obtain a fine powder, which was annealed at 500 °C for 3 h. Subsequently, the calcination process was repeated for an additional 20 min at 800 °C, with intermittent milling after each 3 h interval. Finally, the mixture was heated to the desired temperature at a ramping speed of 10 °C per minute. However, the mechanochemical

technique for synthesizing  $\text{ZnSnO}_3$  has the drawbacks of extended processing times and high energy consumption. Specifically, this method involves mechanical milling or grinding of precursor materials to induce chemical reactions. Although it can yield high-purity  $\text{ZnSnO}_3$ , the need for prolonged milling can be time-consuming and energy-intensive.

Building on the comprehensive review presented earlier, the precise synthesis of diverse ZTO micro-/nanostructures represents a captivating avenue for achieving enhanced performance and investigating the interplay between structure, properties, and performance. Additionally, the utilization of hollow shapes with customizable sizes, morphologies, and compositions has gained significant attention in the fields of sensors and catalysts owing to their exceptional attributes including improved surface area, reduced density, and extensive volume. These remarkable features make them highly desirable for various applications.

## 4. Application

In this section, we present some of the common uses of morphological  $\text{ZnSnO}_3$  nanoparticles in the fields of photocatalysts, sensing, storage, and energy conversion devices.

The ability to design and synthesize  $\text{ZnSnO}_3$  micro/nanostructures with precise control of their morphology, phase, and homogeneous size holds great potential for expanding their applications. Particularly, the crystalline phase plays a crucial role in energy harvesting applications. Moreover, by enhancing the structural or electrical properties of  $\text{ZnSnO}_3$  nanocomposites, they can be utilized in a broad range of applications. Extensive studies exhibited the exceptional potential of  $\text{ZnSnO}_3$  materials in various fields, including gas-sensing, photocatalysis, piezocatalysts, lithium-ion batteries, transparent conductors, and energy storage. Their remarkable electron mobility and electrical conductivity are the main factors responsible for this characteristic.

In this section, we explore the diverse morphology-based applications of  $\text{ZnSnO}_3$  nanocomposites in photocatalysis,



sensing, energy storage and conversion devices, highlighting their significant contributions to these fields.

#### 4.1. Photodegradation of pollutants

Zinc stannate is a ternary metal oxide that holds great potential for photocatalytic and pyroelectric wastewater treatment. Previous studies have established that  $\text{ZnSnO}_3$  exhibits ferroelectric properties, and similar to ferroelectrics with a perovskite structure, display piezoelectric behaviour, making it an important candidate for wastewater treatment.<sup>148–150</sup> Among the six possible crystal structures, LN-type- $\text{ZnSnO}_3$  has been found to possess superior photocatalytic activity due to its piezoelectric property.

The enhanced photocatalytic activity of LN-type  $\text{ZnSnO}_3$  can be attributed to a specific feature of this crystallite orientation. In LN-type  $\text{ZnSnO}_3$ , the displacement of Zn ions next to the z-axis is higher than Sn ions, resulting in spontaneous polarization. The theoretical high spontaneous polarization of LN-type  $\text{ZnSnO}_3$  thin films is estimated to be approximately  $P_r \approx 59 \text{ C cm}^{-2}$ . However, experimental studies have revealed slightly lower values, such as  $P_r \approx 47 \text{ C cm}^{-2}$  at a coercive field of  $E_c \approx 130 \text{ kV cm}^{-1}$  and  $P_r \approx 30 \text{ C cm}^{-2}$  at  $E_c \approx 25 \text{ kV cm}^{-1}$  in single crystal-oriented LN-type  $\text{ZnSnO}_3$  thin films. Vertically aligned LN-type  $\text{ZnSnO}_3$  films exhibit a measured spontaneous polarization of  $P_r \approx 26 \text{ C cm}^{-2}$ .<sup>14</sup>

In the study in ref. 99, the authors employed the sol-gel method to synthesize  $\text{ZnSnO}_3$  catalyst nanoparticles and investigated their behaviour in the process of synergistic pyro- and photo-bi-catalysis. When subjected to UV light and thermal cycles in the range of  $20^\circ\text{C}$  to  $65^\circ\text{C}$ , the  $\text{ZnSnO}_3$  catalyst exhibited a remarkable bi-catalytic dye degradation efficiency of about 98.1%. This efficiency significantly surpassed the photocatalytic degradation (76.8%) and the pyrocatalytic degradation (20.2%). The enhanced photoactivity of  $\text{ZnSnO}_3$  can be attributed to its ability to facilitate faster separation of electron-hole ( $e^-$ - $h^+$ ) pairs, effectively acting as an electron trap. Consequently, the piezo-photocatalyst demonstrated a higher decomposition/breakdown performance than the ordinary photocatalyst. This improvement can be attributed to the synergistic catalytic effect achieved by combining the activity of piezo-phototronics with piezoelectricity, semiconductors, and photonics. This coupling effectively reduces the recombination of  $e^-$ - $h^+$  pairs generated by light and improves their mobility by inducing energy band distortion through applied stresses.

In the study in ref. 151, it was demonstrated that the efficiency of the  $\text{ZnSnO}_3/\text{ZnO}$  composite was significantly higher than that of  $\text{ZnSnO}_3$  alone in the degradation of phenol. The  $\text{ZnSnO}_3/\text{ZnO}$  nanocomposite exhibited an improved photocatalytic breakdown with a reaction rate of  $0.023 \text{ min}^{-1}$  than that of  $0.0168 \text{ min}^{-1}$  for  $\text{ZnSnO}_3$ . This improvement in the reaction rate constant was ascribed to the mixed oxides of the composite with increased surface area of  $\text{ZnSnO}_3/\text{ZnO}$  nanocomposites. Consequently, the effective charge separation of the  $e^-$ - $h^+$  pairs occurred. This study also investigated the surface properties of the  $\text{ZnSnO}_3/\text{ZnO}$  nanocomposite by studying the isoelectric point (IEP) using the pH drift method.

The IEP refers to the pH at which a semiconductor metal oxide carries no net charge, significantly affecting the surface characteristics. The most significant photodegradation of phenol occurred in mildly acidic medium, specifically at a pH of  $\sim 6.4$ . The efficiency of the photocatalytic reaction is directly influenced by the pH of the solution, given that it impacts the formation of hydroxyl radicals and leads to changes in the surface characteristics of the photocatalyst.

Furthermore, in another study in ref. 91,  $\text{ZnSnO}_3$  nanowires were found to exhibit excellent piezophotocatalytic activity, with a rate constant of  $\sim 0.0176 \text{ min}^{-1}$ . This enhanced performance can be attributed to the large surface area, superb arrangement, piezo-potential build-up, and band deforming characteristics of the  $\text{ZnSnO}_3$  nanowires.

The optical-absorption activity of a photocatalyst plays a crucial part in photocatalysis. The authors in ref. 147 synthesized a nanocomposite of S-doped  $\text{ZnSnO}_3$ , which exhibited improved adsorption abilities and a narrow energy gap. The S- $\text{ZnSnO}_3$  material exhibited a significantly high specific surface area, reaching  $80.63 \text{ m}^2 \text{ g}^{-1}$ , which facilitated the effective adsorption of reactants. Additionally, the introduction of S-doping in  $\text{ZnSnO}_3$  led to a rapid reduction in the band gap from 3.7 to 2.4 eV. This significant reduction in the band gap resulted in a highly effective photocatalyst for the decomposition of RhB, achieving an efficiency of almost 90%, which is substantially higher than when using pure  $\text{ZnSnO}_3$ .

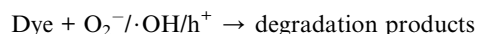
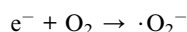
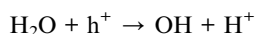
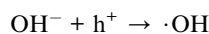
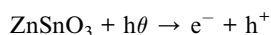
Furthermore, the manipulation of the surface morphologies of  $\text{ZnSnO}_3$  also improves its photocatalytic capabilities by reducing the photon-induced recombination  $e^-$ - $h^+$  pairs and enhancing their mobility. In the in ref. 116, the authors designed double-shelled  $\text{ZnSnO}_3$  hollow-cubes to enhance the photocatalytic breakdown of antibiotic wastewater. These double-shelled  $\text{ZnSnO}_3$  hollow cubes provide a high surface area, offer large reactive sites, and allow multi-scattering of incident light, thereby enhancing the photocatalytic performance. The trapping experiments conducted in the study detected the primary reactive species in the photo breakdown activity as  $\cdot\text{OH}$  and  $h^+$ . These active species successfully encourage the breakdown of organic dyes under exposure to simulated sunlight.

Hollow  $\text{ZnSnO}_3$  nanospheres/rGO nanocomposites have been proven to be effective photocatalysts for degrading metronidazole in aqueous solutions. These nanocomposites demonstrated remarkable photodegradation capabilities, particularly under visible light irradiation for a duration of 180 min. The hollow  $\text{ZnSnO}_3$  nanospheres achieved a degradation efficiency of 42.1%, while the  $\text{ZnSnO}_3/\text{rGO}$  nanocomposites exhibited an even higher efficiency of 72.5%. In contrast, the pure rGO and  $\text{ZnSnO}_3$  with rGO physical mixtures displayed a very low performance in terms of photodegradation. The outstanding performance of  $\text{ZnSnO}_3/\text{rGO}$  can be related to several factors. Firstly, the hybrid nanocomposite possesses higher adsorption effectiveness for the target dyes. Additionally, it exhibits enhanced absorption of visible light due to the presence of rGO. Finally, the distinctive electronic system of the  $\text{ZnSnO}_3/\text{rGO}$  hybrid nanocomposite



contributes to its enhanced photocatalytic properties. This finding was reported in ref. 132.

The probable reaction mechanism pathway in the breakdown of dye can be outlined as follows:<sup>116,147</sup>



When light of frequency  $\nu$  equaling the  $E_g$  of metal oxide semiconductors illuminates them, the electrons ( $e^-$ ) in the semiconductors become excited and transition from the VB to the CB. This process generates holes ( $h^+$ ) in the VB, resulting in the formation of  $e^-h^+$  pairs. These  $h^+$  and  $e^-$  can engage in reactions with  $\text{H}_2\text{O}$  molecules,  $\text{OH}^-$  ions, and dissolved  $\text{O}_2$ , leading to the creation of  $\cdot\text{OH}$  radicals and  $\cdot\text{O}_2^-$  radicals, respectively. Conversely, the organic dye can also undergo photosensitization. When a photosensitized organic dye undergoes dissociation in water, it can react with  $\cdot\text{OH}$  radicals and  $\cdot\text{O}_2^-$  radicals to produce  $\text{H}_2\text{O}$  and  $\text{CO}_2$ .

After band-gap excitation and the generation of free charge carriers (electrons and holes), various mechanisms of de-excitation occur, as shown in Fig. 17a and b.<sup>116</sup> Together with the desired redox reactions (pathways 1 and 2), several recombination processes compete with each other, hindering the successful transfer of the carriers to acceptor molecules at the surface.

These competing recombination processes include the recombination of carriers with their oppositely charged

counterparts trapped on the surface (route 3) and the recombination of two carriers in the bulk of the semiconductor (pathway 4). Both of these mechanisms contribute to a decrease in the efficiency of the photocatalytic reaction. The rate of recombination is influenced by several factors, including the mobility and trapping behaviour of the charge carriers, the density of defects in the semiconductor lattice, and the presence of an interface with a secondary material, which acts as a sink for either electrons or holes. These factors collectively affect the recombination rates, and consequently impact the overall effectiveness of the photocatalytic process (Table 4).<sup>152</sup>

## 4.2. Gas sensors

Extensive research has been conducted in recent decades on gas-sensing materials due to their crucial role in detecting hazardous gases. This research has significant implications in industry, environmental monitoring, agriculture, medical diagnosis, military, human health, and aerospace.<sup>161–164</sup> Gas sensors based on n-type semiconductors have been employed to detect inflammable and hazardous gases such as  $\text{H}_2$ ,  $\text{H}_2\text{S}$ ,  $\text{CH}_4$ ,  $\text{C}_2\text{H}_5\text{OH}$ , and  $\text{CO}$ . However, a major limitation of these sensors is their lack of selectivity.

Thus, to address this limitation, the n-type  $\text{ZnSnO}_3$  semiconductor has emerged as a promising solution for enhancing the selectivity of gas sensors towards various gases. The utilization of the sensitive material has been greatly improved by the substantial increase in the specific surface area of  $\text{ZnSnO}_3$ . This enhanced specific surface area of the semiconductor allows for better capacity to adsorb gas molecules. Additionally, the double-shelled hollow morphology with mesoporous features not only facilitates the adsorption of gas molecules on the surface of the sensitive/reactive material but also allows for deep penetration of gas molecules inside the hollow spheres. This unique structure enables the rapid diffusion of gas

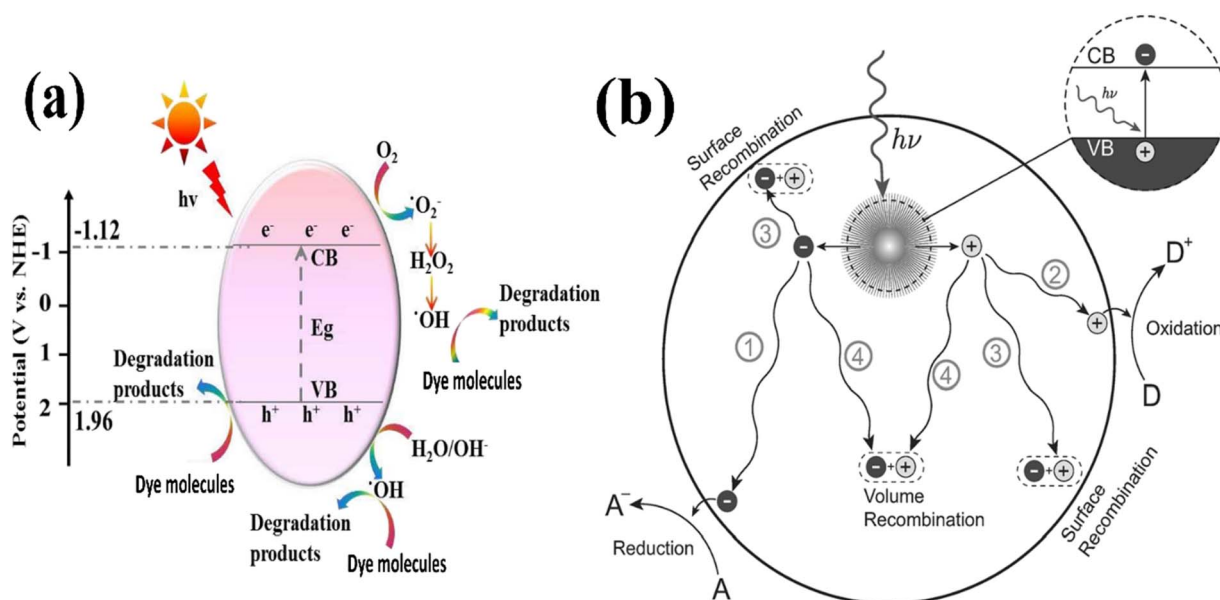


Fig. 17 Illustration of (a) photocatalytic mechanism. (b) Photoinduced formation of an electron hole pair in a semiconductor.<sup>148,217</sup>



Table 4 Photodegradation of pollutants based on ZnSnO<sub>3</sub> materials and their performance

| Materials                                                             | Morphology         | Light source  | Time | Dye           | Efficiency (%)    | Ref. |
|-----------------------------------------------------------------------|--------------------|---------------|------|---------------|-------------------|------|
| ZnSnO <sub>3</sub> /rGO                                               | Nanosheets         | Visible-light | 35   | MB            | 85                | 76   |
| ZnSnO <sub>3</sub>                                                    | Nanocubes          | UV            | 90   | Phenol        | 72.89             | 99   |
|                                                                       |                    |               | 80   | RhB           | 98.1 <sup>a</sup> |      |
| S-ZnSnO <sub>3</sub>                                                  | Hollow cubes       | Visible-light | 100  | RhB           | 76.8              | 147  |
| ZnSnO <sub>3</sub> /rGO/MoS <sub>2</sub>                              | Nanosheets         | UV            | 100  | MB            | 90                | 153  |
|                                                                       |                    |               | 100  | RhB           | 86                |      |
| ZnSnO <sub>3</sub>                                                    | Hollow cubes       | Xe            | 100  | MB            | 78                | 116  |
|                                                                       |                    |               |      | CIP           | 96.52             |      |
| ZnSnO <sub>3</sub> /rGO                                               | 3D folded          | Visible       | 120  | CIP           | 85.9              | 154  |
| ZnSnO <sub>3</sub>                                                    | Nanostructures     | UV            | 120  | Acid brown 14 | 100               | 103  |
| ZnSnO <sub>3</sub> /rGO                                               | Hollow nanospheres | Visible light | 180  | Metronidazole | 92                | 132  |
| ZnSnO <sub>3</sub> /ZnO                                               | Nanocubes          | UV            | 60   | Phenol        | 72.5              | 132  |
| AgI/ZnSn(OH) <sub>6</sub>                                             | Nanocubes          | Visible light | 40   | Phenol        | 47                | 151  |
|                                                                       |                    |               | 60   | RhB           | 100               | 155  |
| ZnSnO <sub>3</sub>                                                    | Nanourchins        | UV            | 60   | TOC           | 91                | 75   |
|                                                                       |                    |               | 90   | MO            | 98                |      |
| ZnSnO <sub>3</sub> /g-C <sub>3</sub> N <sub>4</sub>                   | Nanosheets         | Visible       | 120  | 2,5-DCP       | 95                | 156  |
| ZnSnO <sub>3</sub> /g-C <sub>3</sub> N <sub>4</sub>                   | Nanocubes          | Visible       | 120  | Tetracycline  | 85                | 157  |
| Cu <sub>3</sub> P/ZnSnO <sub>3</sub> /g-C <sub>3</sub> N <sub>4</sub> | Nanocubes          | Xe lamp       | 60   | Tetracycline  | 90.8              | 158  |
| CoP/ZnSnO <sub>3</sub>                                                | Nanocubes          | Visible       | 60   | Tetracycline  | 98.45             | 159  |
| CDs-ZnSnO <sub>3</sub>                                                | Nanocubes          | Visible       | 60   | Tetracycline  | 96.44             | 144  |
| NDs/ZnSnO <sub>3</sub>                                                | Nanocubes          | Visible       | 60   | Tetracycline  | 81.76             | 144  |
| Ag-ZnSnO <sub>3</sub>                                                 | Hollow nanocubes   | UV            | 120  | Tetracycline  | 90                | 160  |
|                                                                       |                    |               | 240  | Caffeine      | 68 <sup>b</sup>   | 143  |
|                                                                       |                    |               |      |               | 100 <sup>c</sup>  |      |

<sup>a</sup> Pyro-/photo-bi-catalysis. <sup>b</sup> pH = 6.5. <sup>c</sup> pH = 7.5.

molecules and oxygen, thereby improving the overall performance of the gas sensor.

ZnSnO<sub>3</sub> is widely recognized as an n-type semiconductor, primarily governed by the behaviour of its conduction electrons. When ZnSnO<sub>3</sub>-based sensors come into contact with air, oxygen molecules tend to adsorb on the sample surface, effectively capturing free electrons from the CB. This process leads to the formation of ionized oxygen species, including O<sub>2</sub><sup>-</sup>, O<sup>-</sup> and O<sup>2-</sup>. Consequently, the consumption of electrons results in the creation of a substantial electron depletion layer on the surface of the sample. This electron depletion layer significantly increases the resistance of the sensor, thereby influencing its overall performance.

ZnSnO<sub>3</sub> has been extensively studied as a highly effective material for detecting gases that are reducing and combustible. A gas sensor fabricated using polyhedral ZnSnO<sub>3</sub> demonstrated exceptional sensitivity, quick recovery, and reliable results for the detection of HCHO gas. Among the various polyhedral shapes, the 14-faceted ZnSnO<sub>3</sub>, consisting of 6 {100} facets and 8 {111} facets, exhibited significantly greater sensitivity compared to the octahedral shape and commercially available powder. This improved sensitivity is attributed to the larger exposure of the active {100} facets on the 14-faceted ZnSnO<sub>3</sub>. The increased number of active {100} facets in the 14-faceted ZnSnO<sub>3</sub> led to a higher specific surface area, providing more reactive sites for ZnSnO<sub>3</sub> to bind gas molecules. Consequently, this results in higher sensitivity. Considering these findings, it is promising to

explore the potential of using cubic ZnSnO<sub>3</sub> entirely exposed with {100} facets for gas sensing applications.<sup>32</sup>

The researchers in ref. 165 have successfully developed hollow ZnSnO<sub>3</sub> nanocages, which showed remarkable gas sensing properties, particularly when detecting formaldehyde gas. These nanocages demonstrated a high gas response, together with short response and recovery times, superior selectivity, and good repeatability and stability, even at low concentrations of HCHO. This study confirmed that the highest response, with a value of 57.6, was achieved when exposed to 50 parts per million (ppm) of HCHO gas at 350 °C. The response time for the nanocages was measured to be 3 s, while the recovery time was 5 s. Alternatively, the sensitivity of these nanocages towards other gases such as C<sub>2</sub>H<sub>5</sub>OH, (CH<sub>3</sub>)<sub>2</sub>CO, CH<sub>3</sub>OH, and NH<sub>3</sub> was significantly lower compared to formaldehyde gas.

To enhance the selectivity of sensors, the researchers in ref. 126 investigated the use of hollow cubic nanoparticles of zinc stannate with a surface area of 22.6319 m<sup>2</sup> g<sup>-1</sup>, which were synthesized through the hydrothermal process. The sensor fabricated using these samples exhibited significant improvements in selectivity. When exposed to 500 ppm of butane gas at an operating temperature of 380 °C, the response of the hollow microspheres was measured to be 5.79, whereas the response of the solid microspheres was 3.92. Moreover, the response time and recovery time of the hollow ZnSnO<sub>3</sub> microspheres were recorded to be 0.3 s and 0.65 s, while that of the solid ZnSnO<sub>3</sub> microspheres was 1.3 s and 13.7 s, respectively. These superior



features of the sensor were ascribed to the larger specific surface area of the hollow particles in comparison to the solid particles. This enabled better access to the surface, leading to an increased ability to adsorb gas molecules.

Jiang *et al.* employed a hydrothermal technique to prepare perovskite-type  $\text{ZnSnO}_3$  and conducted gas sensitivity experiments.<sup>166</sup> The results obtained from calculations revealed that acetone molecules strongly interacted with pre-adsorbed  $\text{O}_2^-$  and  $\text{O}^-$  on the  $\text{ZnSnO}_3$  (001) surface, leading to charge transfer and resistance changes in the material. These observed phenomena form the basis for utilizing  $\text{ZnSnO}_3$  as a sensing material for detecting acetone. The experimental findings demonstrated that the perovskite-type  $\text{ZnSnO}_3$  exhibited a sheet like morphology with micro-holes on its surface, which contributed to enhanced sensitivity towards  $(\text{CH}_3)_2\text{CO}$  gas. Through gas sensing tests, the ideal performance temperature was determined to be 350 °C, with a response time of 4 s and recovery time of 27 s. The sensitivity of the  $\text{ZnSnO}_3$  nanosheets produced in this study towards 100 ppm and 10 ppm of  $(\text{CH}_3)_2\text{CO}$  gas was found to be 125.444 and 8.37, respectively. Furthermore, the sensor based on the sheet-like  $\text{ZnSnO}_3$  maintained 89.64% and 94.74% of its original sensitivity towards 10/100 ppm acetone gas, respectively, after undergoing a stability and repeatability test for five weeks.

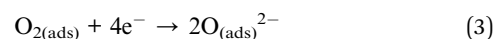
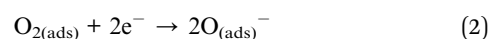
One approach to enhance the selectivity towards specific gases involves doping or modifying the gas-sensitive material through the use of another substance and creating composites. This was demonstrated in ref. 167, where  $\text{Ag}/\text{ZnSnO}_3$  structures exhibited sensitivity towards acetone. In this study, it was found that the response to 100 ppm of acetone was 30, with a response time of 2 s and recovery time of 3 s. The optimal operating temperature was determined to be 280 °C. In another study, ref. 168, focused on increasing the sensitivity towards ethanol vapor by doping with  $\text{NiO}$ . The sensitivity to 20 ppm of  $\text{C}_2\text{H}_5\text{OH}$  was measured at 23.95, with a response time of 56 s and recovery time of 48 s. The optimal operating temperature for this system was found to be 160 °C. In ref. 169, the authors synthesized zinc stannate nanofibers with a surface area of  $88.4 \text{ m}^2 \text{ g}^{-1}$  using the electrospinning method. The sensitivity towards acetone was enhanced by doping with  $\text{Au}/\text{In}-\text{ZnSnO}_3$ . The response to 50 ppm of ethanol was recorded at 19.3, with a response time of 10 s and recovery time of 13 s at 200 °C.

The gas-sensing properties of conventional  $\text{ZnSnO}_3$  nanoparticles are often unsatisfactory for sensor applications due to their high operating temperatures. These sensors typically require temperatures exceeding 200 °C, which limits their practical use. At lower temperatures, the thermal energy is insufficient to overcome the activation energy barrier, resulting in inadequate reactions with sensitive materials. Simultaneously, as the operating temperature increases, gas molecules gain higher thermal energy, enabling them to surpass the activation energy barrier. This leads to enhanced reactions among gas molecules and adsorbed oxygen, as well as enhanced ion adsorption on the surface of the semiconductor material.

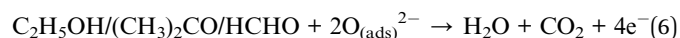
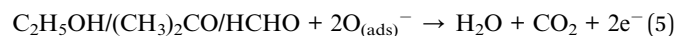
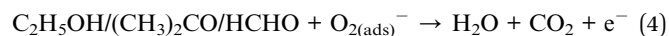
Thus, to address this issue and achieve high-performance gas sensors with reduced working temperatures, various

modifications have been applied, such as preparing nano-materials with different morphologies and doping. In ref. 170, the authors synthesized 3D flower-like  $\text{ZnSnO}_3$  using the co-precipitation method and improved the  $\text{C}_2\text{H}_5\text{OH}$  sensing properties at RT by incorporating nano- $\text{TiO}_2$ . This modification significantly enhanced the ethanol response at room temperature, reaching a maximum value of 36.6. The response and recovery time for a low concentration of 10 ppm  $\text{C}_2\text{H}_5\text{OH}$  were recorded to be 4 s and 17 s, respectively. The authors attributed the formation of the 3D flower-like structure to the presence of the stabilizer and pH balancer triethanolamine (TEOA) during the synthesis of  $\text{ZnSnO}_3$ .

$\text{ZnSnO}_3$  is a conventional n-type metal-oxide semiconductor and its gas sensing mechanism operates through surface control, relying on changes in conductivity during the gas-sensing reaction. Upon exposure to air,  $\text{ZnSnO}_3$  nanocubes adsorb oxygen ions on their surface, as follows:<sup>171</sup>



When the  $\text{ZnSnO}_3$  nanocubes are exposed to  $\text{C}_2\text{H}_5\text{OH}/(\text{CH}_3)_2\text{CO}/\text{HCHO}$ , the  $\text{C}_2\text{H}_5\text{OH}/(\text{CH}_3)_2\text{CO}/\text{HCHO}$  molecules will interact with adsorbed oxygen ions as follows:



Eqn (1) to (3) describe the process in which electrons are extracted from the CB of  $\text{ZnSnO}_3$ , leading to an increase in resistance. Alternatively, eqn (4) to (6) depict the process where electrons are returned to the CB of  $\text{ZnSnO}_3$ , resulting in a decrease in resistance. This gas-sensing mechanism highlights that the quantity of surface-adsorbed oxygen ions plays a crucial role in the gas-sensing characteristics of  $\text{ZnSnO}_3$ . As electrons are consumed in the surface region of the sensitive material, a depletion layer is formed, leading to an increase in the resistance of the material, as shown in Fig. 18.<sup>170</sup>

#### 4.3. Anode for lithium-ion batteries

Lithium-ion batteries (LIBs) have revolutionized the energy storage landscape and become the backbone of portable electronic devices, electric vehicles, and grid storage applications. The remarkable combination of high energy density, long cycle life, and low self-discharge rate has made Li-ion batteries the preferred choice in various industries.<sup>186–188</sup> Over the years, extensive research efforts have been directed towards enhancing the performance and safety of Li-ion batteries, driven by the growing demand for sustainable and efficient energy storage solutions. In this case,  $\text{ZnSnO}_3$  possesses a unique crystal structure and a wide bandgap, making it an intriguing material for Li-ion battery applications. When used



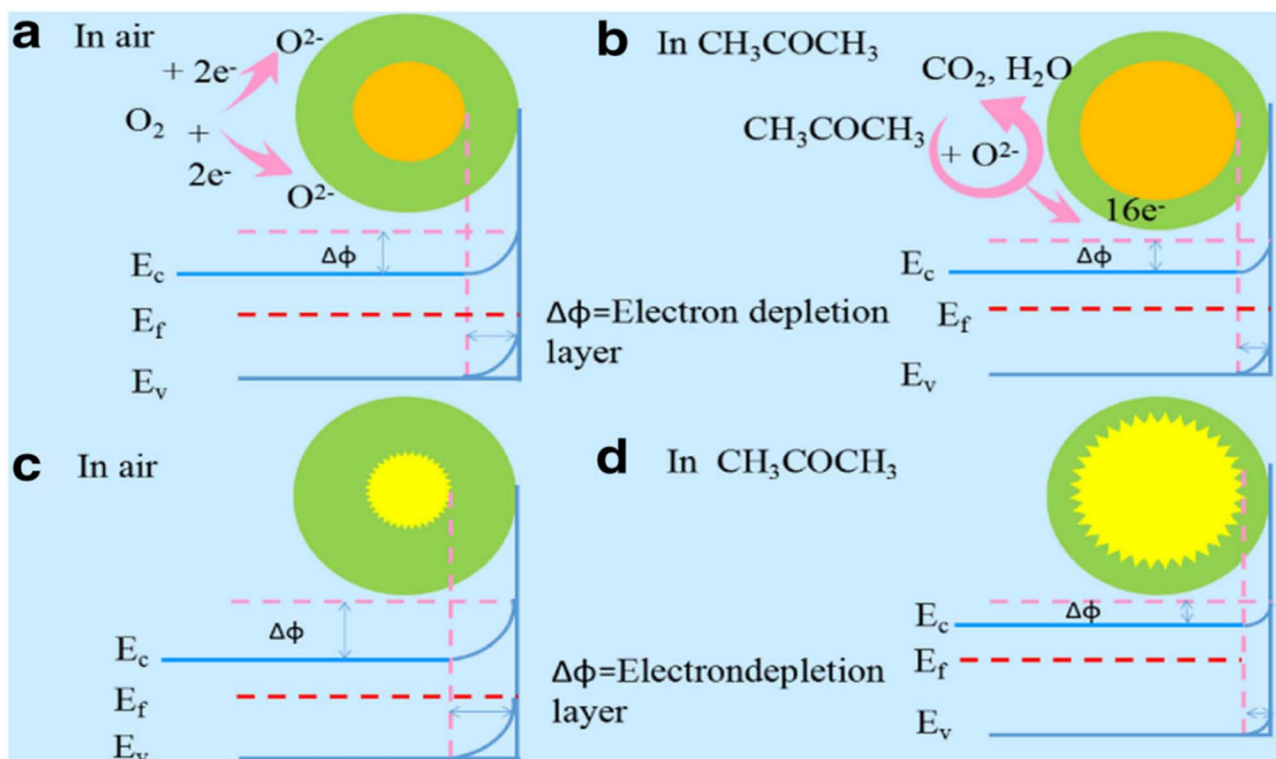


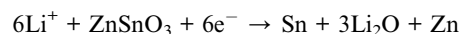
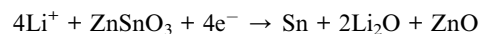
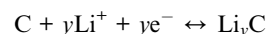
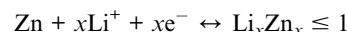
Fig. 18 Schematic band diagrams of ZnSnO<sub>3</sub> hollow microspheres exposed to (a) air and (b) CH<sub>3</sub>COCH<sub>3</sub>, and ZnSnO<sub>3</sub>/SnO<sub>2</sub> hollow urchin gas sensors exposed to (c) air and (d) CH<sub>3</sub>COCH<sub>3</sub>.<sup>218</sup>

as an electrode material, ZnSnO<sub>3</sub> exhibits an exceptional theoretical capacity of 1320 mA h g<sup>-1</sup>, surpassing that of many traditional alternatives.<sup>71</sup> This high capacity implies that ZnSnO<sub>3</sub> has potential to store more lithium ions, resulting in increased energy storage capabilities (Table 5).

Although the various structures of ZnSnO<sub>3</sub> demonstrate appealing characteristics in LIBs, this material still faces notable changes in volume during repeated charging and discharging cycles. Consequently, its rate performance is compromised and it experiences considerable capacity loss, impeding its practical use in real-world applications.<sup>188,189</sup> Thus, to address the aforementioned limitations of ZnSnO<sub>3</sub>, scientists have implemented diverse approaches, including the creation of micro-nano structures, incorporating transition metals through modification, and combining the material with carbon substances.<sup>190–193</sup>

A reported study demonstrated that a sandwich-like C/ZnSnO<sub>3</sub> composite exhibited comparatively higher reversible capacity even after long-term cycling tests.<sup>194</sup> Binary ZnSnO<sub>3</sub> nanocomposites combine two practical materials, displaying an improved synergistic performance. The integration of these elements results in an enhancement of their individual properties, including improved electrical and ionic conductivity, enhanced electrochemical reactivity, and increased mechanical stability. The sandwich-like C/ZnSnO<sub>3</sub> nanocomposite displayed a reversible capacity of 1107 mA h g<sup>-1</sup> under a current density of 100 mA g<sup>-1</sup>, with a coulombic output of ~98.5% after the 200th cycle. This performance was better

than that of sheet-like ZnSnO<sub>3</sub>, which showed a coulombic output of ~96.2% after 100 cycles. Fig. 19 illustrates the synthesis procedure of the sandwich-like ZnSnO<sub>3</sub>@C. The charge-discharge data of the sandwich-like ZnSnO<sub>3</sub>@C and sheet-like ZnSnO<sub>3</sub> at a current density of 100 mA g<sup>-1</sup> and the cycling performance of the nanocomposite are shown in Fig. 20a–d. The complete electrochemical reaction pathway can be outlined as follows:<sup>71,194</sup>

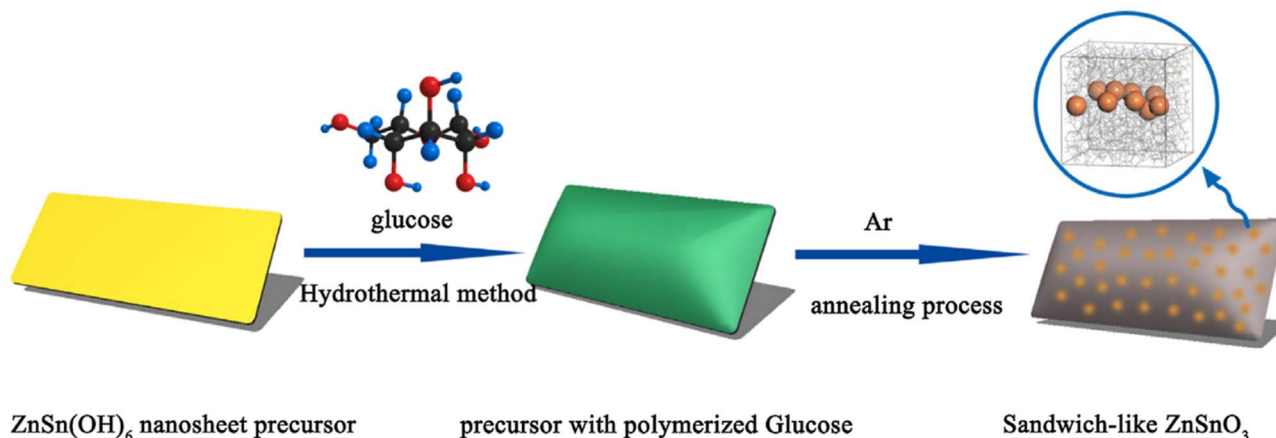


Overall, the unique nanostructure of the sandwich-like C/ZnSnO<sub>3</sub> composite, with its interconnected carbon nanoplates, encapsulated ZnSnO<sub>3</sub> nanoparticles, and high specific surface area (106 m<sup>2</sup> g<sup>-1</sup>), enables enhanced Li storage activity by facilitating efficient charge transfer, preventing material detachment, and accommodating volume changes during electrochemical reactions.<sup>194</sup>

Amorphous ZnSnO<sub>3</sub> double-shell hollow microcubes (D-ZnSnO<sub>3</sub>) possessed an average edge length of 1.6 μm and

Table 5 Gas sensors based on ZnSnO<sub>3</sub> and their performance

| Sensing materials                    | Morphology                   | Surface area (m <sup>2</sup> g <sup>-1</sup> ) | Target             | Concentration (ppm) | Operating temp. (°C) | Response (R <sub>a</sub> /R <sub>g</sub> ) | Response-recovery time (s) | Ref. |
|--------------------------------------|------------------------------|------------------------------------------------|--------------------|---------------------|----------------------|--------------------------------------------|----------------------------|------|
| ZnSnO <sub>3</sub>                   | Hollow microspheres          | 62.50                                          | Ethanol            | 500                 | 320                  | 144.3                                      | 7, 8                       | 90   |
|                                      | Solid microspheres           | 40.43                                          |                    |                     | 260                  | 80                                         | 4, 7                       |      |
| CeO <sub>2</sub> /ZnSnO <sub>3</sub> | Hollow microspheres          | —                                              | Ethanol            | 100                 | 200                  | 219.2                                      | 12, 22                     | 172  |
| CuO/ZnSnO <sub>3</sub>               | Hollow microspheres          | —                                              | Ethanol            | 100                 | 160                  | 131                                        | 13, 8                      | 173  |
| Ag-ZnSnO <sub>3</sub>                | Nanocubes                    | —                                              | Ethanol            | 100                 | 200                  | 83.96                                      | 1, 50                      | 174  |
| ZnSnO <sub>3</sub>                   | Hollow microspheres          | 30.21                                          | Ethanol            | 100                 | 280                  | 24                                         | —                          | 70   |
| NiO/ZnSnO <sub>3</sub>               | Microsphere-decorated fibers | 60                                             | Ethanol            | 20                  | 160                  | 23.95                                      | 56, 48                     | 168  |
| ZnSnO <sub>3</sub>                   | Hollow spheres               | —                                              | Ethanol            | 50                  | 270                  | 27.8                                       | 0.9, 2.2                   | 127  |
| ZnSnO <sub>3</sub>                   | 2D nanosheets                | 85.1                                           | HCHO               | 10                  | 100                  | 57                                         | 36, 42                     | 175  |
|                                      | Nanocubes                    | 59.7                                           | Ethanol            | 200                 | 200                  | 20.3                                       | 41, 105                    |      |
| ZnSnO <sub>3</sub>                   | 3D flower-like               | —                                              | Ethanol            | 1000                | RT(25)               | 36.6                                       | 2, 14                      | 170  |
| ZnSnO <sub>3</sub>                   | Hollow nanocages             | 76.4                                           | HCHO               | 50                  | 350                  | 57.6                                       | 3, 5                       | 165  |
|                                      |                              |                                                | Ethanol            |                     |                      | 34.5                                       |                            |      |
|                                      |                              |                                                | Acetone            |                     |                      | 23.6                                       |                            |      |
| ZnSnO <sub>3</sub>                   | Nanosheets                   | 57.86                                          | Ethanol            | 50                  | 320                  | 33                                         | 0.36, 9                    | 94   |
|                                      |                              |                                                | Acetone            |                     | 440                  | 41                                         | 0.18, 2.6                  |      |
| ZnSnO <sub>3</sub>                   | Nanocubes                    | 71                                             | TEA                | 100                 | 350                  | 57.5                                       | 41, 040                    | 72   |
| ZnSnO <sub>3</sub>                   | Hierarchical                 | 67.8                                           | HCHO               | 10                  | 300                  | 34.1                                       | 132, 15                    | 171  |
| ZnSnO <sub>3</sub>                   | Hollow cubes                 | 98                                             | HCHO               | 100                 | 220                  | 37.2                                       | 1, 59                      | 124  |
| ZnSnO <sub>3</sub>                   | Hollow cubes                 | —                                              | HCHO               | 200                 | 270                  | 22.5                                       | —                          | 136  |
| ZnSnO <sub>3</sub>                   | Nanocubes                    | 70.001                                         | HCHO               | 50                  | 210                  | 21.2                                       | —                          | 137  |
| ZnSnO <sub>3</sub>                   | Porous nanostructures        | —                                              | HCHO               | 100                 | 240                  | 13.5                                       | 1, 10                      | 176  |
| ZnSnO <sub>3</sub>                   | Porous cubes                 | 33.7                                           | Toluene            | 100                 | 300                  | 16.56                                      | 2, 11                      | 177  |
|                                      |                              |                                                | HCHO               |                     |                      | 36.81                                      | 3, 12                      |      |
| Zn <sub>2</sub> SnO <sub>4</sub>     | Hierarchical                 | 23.7                                           | Toluene            | 100                 | 280                  | 25.2                                       | 1, 3.5                     | 178  |
|                                      | Spherical                    | 20.6                                           |                    |                     |                      | 19.2                                       |                            |      |
|                                      | Cubic                        | 9.1                                            |                    |                     |                      | 11.7                                       |                            |      |
| ZnSnO <sub>3</sub>                   | Sheet-like                   | —                                              | Acetone            | 100                 | 350                  | 125.44                                     | 7, 27                      | 166  |
| ZnSnO <sub>3</sub> /rGO              | Hierarchical                 | 33.14                                          | Acetone            | 50                  | 350                  | 80.37                                      | 4, 27                      | 179  |
| Ag/ZnSnO <sub>3</sub>                | Hollow nanocubes             | 19.1                                           | Acetone            | 100                 | 280                  | 30                                         | 2, 3                       | 167  |
| ZnSnO <sub>3</sub> /SnO <sub>2</sub> | Concave microcubes           | 69.7                                           | Acetone            | 50                  | 260                  | 19.1                                       | 5.6                        | 180  |
| ZnSnO <sub>3</sub>                   | Hollow polyhedrons           | —                                              | Acetone            | 50                  | 240                  | 12.48                                      | 17, 10                     | 181  |
| Au/In-ZnSnO <sub>3</sub>             | Nanofibers                   | 88.4                                           | Acetone            | 50                  | 200                  | 19.3                                       | 10, 13                     | 169  |
| ZnSnO <sub>3</sub>                   | Nanocages                    | —                                              | H <sub>2</sub> S   | 100                 | 310                  | 29.4                                       | 20, 50                     | 128  |
| ZnSnO <sub>3</sub> -C1               | Hollow cubes                 | —                                              | H <sub>2</sub> S   | 100                 | 335                  | 14.18                                      | 6, 22                      | 182  |
| ZnSnO <sub>3</sub>                   | Hollow hexahedron            | 152                                            | CO <sub>2</sub>    | 400                 | RT                   | 4, 65                                      | 73.2, 79.5                 | 183  |
| PdO-ZnSnO <sub>3</sub>               | Hollow microspheres          | —                                              | <i>n</i> -Propanol | 100                 | 140                  | 30.8                                       | 1, 25                      | 184  |
| ZnSnO <sub>3</sub>                   | Hollow spheres               | 53.91                                          | <i>n</i> -Butanol  | 50                  | 200                  | 7.1                                        | 2, 40                      | 185  |
| ZnSnO <sub>3</sub>                   | Hollow microspheres          | 22.632                                         | Butane             | 500                 | 380                  | 5.79                                       | 0.3, 0.65                  | 126  |
|                                      | Solid microspheres           | —                                              |                    |                     |                      | 3.92                                       | 1.3, 13.7                  |      |

Fig. 19 Schematic illustration depicting the synthesis of sandwich-like ZnSnO<sub>3</sub>@C.<sup>194</sup>

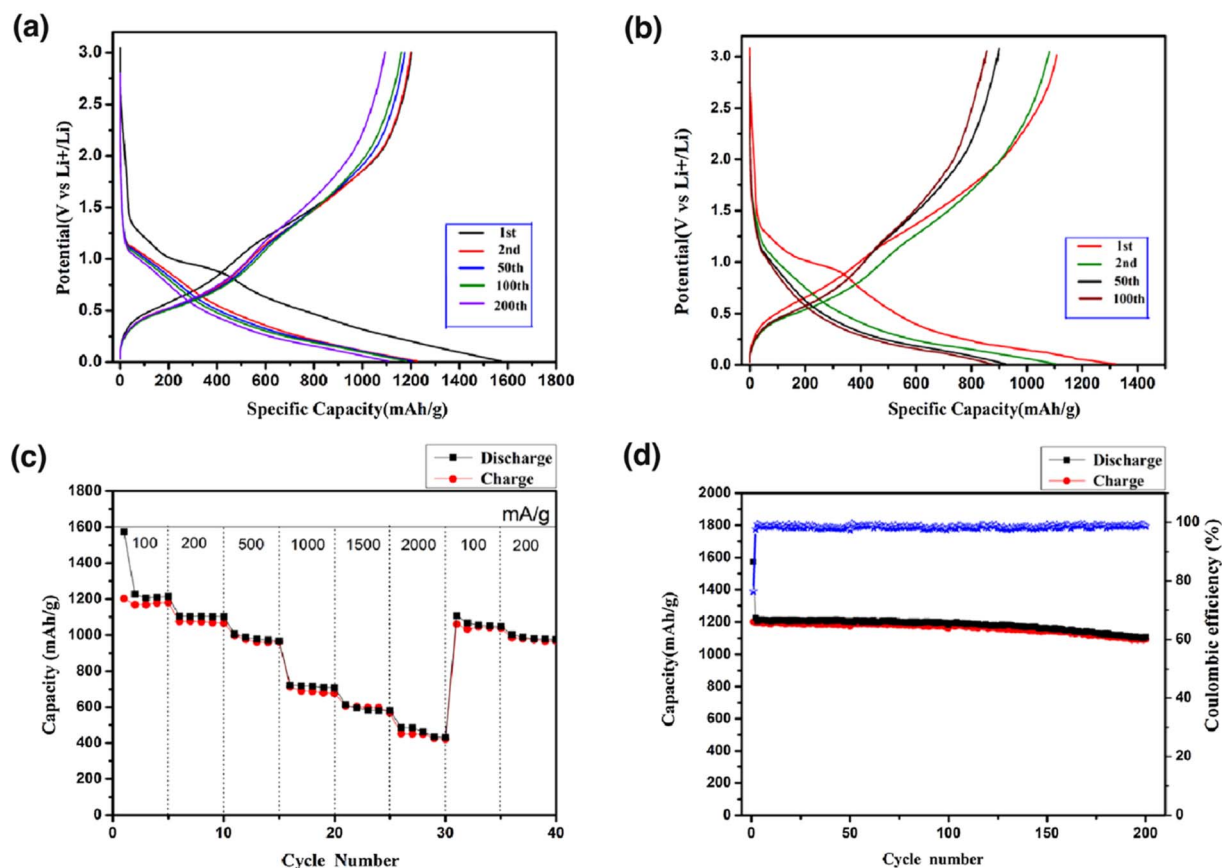


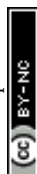
Fig. 20 Charge–discharge profiles of (a) sandwich-like  $\text{ZnSnO}_3\text{@C}$  and (b) sheet-like  $\text{ZnSnO}_3$  at  $100 \text{ mA g}^{-1}$ . (c) Cycling activity at various current densities. (d) Cycling activity at  $100 \text{ mA g}^{-1}$  of the sandwich-like  $\text{ZnSnO}_3\text{@C}$ .<sup>194</sup>

displayed favourable electrochemical characteristics when utilized as anode materials.<sup>195</sup> Compared to their  $\text{ZnSnO}_3$  yolk-shell ( $\text{Y-ZnSnO}_3$ ) counterparts, the  $\text{D-ZnSnO}_3$  anodes exhibited a superior performance. They maintained a high coulombic efficiency of 99% at a current density of  $100 \text{ mA g}^{-1}$  and retained a significant reversible capacity of  $741 \text{ mA h g}^{-1}$  after 50 cycles. The exceptional electrochemical properties of  $\text{D-ZnSnO}_3$  originate from its amorphous nature and unique box-in-box hollow architecture. The amorphous structure of  $\text{D-ZnSnO}_3$  contributed to its isotropy, effectively relieving mechanical strain and preserving the electrode integrity throughout cycling. Furthermore, the hollow box-in-box architecture provided ample space to accommodate the substantial volume changes associated with the insertion and extraction of  $\text{Li}^+$  ions.<sup>195</sup>

Lu *et al.* successfully fabricated amorphous isotropic  $\text{ZnSnO}_3$  nanocubes with a well-defined mesoporous morphology, which were coated with a uniform thin carbon layer, forming core-shell nanoparticles that were interconnected to create a cohesive framework.<sup>196</sup> Comparative electrochemical analysis of the amorphous  $\text{ZnSnO}_3\text{@C}$  and crystalline  $\text{ZnSnO}_3\text{@C}$  revealed that the former exhibited a superior performance, including a higher specific capacity, superior cyclic stability, and favourable rate capability.

Although the capacity and cyclic stability of the crystalline  $\text{ZnSnO}_3\text{@C}$ , with excellent crystallinity, showed a slight increase, reaching a capacity of  $791 \text{ mA h g}^{-1}$  after 100 cycles, the amorphous  $\text{ZnSnO}_3\text{@C}$  demonstrates an improved reversible specific capacity and a more stable capacity retention. The amorphous material achieved reversible capacities exceeding  $1020 \text{ mA h g}^{-1}$  during the initial 20 cycles, with a marginal increase to  $1060 \text{ mA h g}^{-1}$  after 100 cycles. This enhancement was attributed to the improved accessibility of lithium ions in the structure. The uninterrupted carbon nanocoating, interconnected porous structure, well-established mesoporosity, and homogeneous amorphous composition all significantly contributed and played essential roles in the accommodation of mechanical strain, buffering of volume changes, and maintenance the structural integrity of  $\text{ZnSnO}_3$  to achieve a stable cycling performance.

Additionally, this unique structure facilitates rapid electron and Li-ion transfer, thereby enhancing the application of active materials and enabling high-reversible cathode performances.<sup>196</sup> The utilization of  $\text{ZnO@ZnSnO}_3$  quantum dot (QD) core-shell nanorod (NR) arrays as anodes led to significant improvements in their electrochemical performance.<sup>197</sup> The incorporation of a highly conductive  $\text{ZnSnO}_3$  QD shell on the  $\text{ZnO@ZnSnO}_3$  QD core-shell NRs resulted in larger free spaces, which were retained after extended cycling at high current



densities. This indicates that the unique arrangement of the ZnO@ZnSnO<sub>3</sub> QD core-shell NRs successfully accommodated the volume changes during cycling.

The use of a binder-free electrode offers several advantages, such as establishing good contact with the electrolyte and maintaining a short Li-ion diffusion path within the ZnO@ZnSnO<sub>3</sub> QD electrode. Moreover, after 450 cycles at a current density of 1000 mA g<sup>-1</sup>, there were still more free spaces observed among the ZnO@ZnSnO<sub>3</sub> QDs. This suggests that the barrier posed by significant volume expansion can be overcome when the surface of ZnO nanorods is enhanced with ZnSnO<sub>3</sub> QD shells.<sup>197</sup> The incorporation of ZnSnO<sub>3</sub> hollow cubes in flexible reduced graphene oxide (rGO) sheets, resulting in a porous architecture, was proven to be advantageous in lithium-ion battery applications.<sup>146</sup> This design effectively mitigated the issues associated with significant volume enlargement throughout repetitive charge-discharge cycles. Additionally, the 3D networks created by the ZnSnO<sub>3</sub> hollow cubes encased in rGO sheets facilitated the rapid transport of lithium ions and electrons.

As a result of these features, the ZnSnO<sub>3</sub>-graphene assemblies (ZGAs) exhibited an improved rate capacity, achieving a high specific capacity of 552.6 mA h g<sup>-1</sup> at a current density of 1200 mA g<sup>-1</sup>. Furthermore, ZGAs demonstrated excellent cycling stability, maintaining a specific capacity of 745.4 mA h g<sup>-1</sup> after 100 cycles at a current density of 100 mA g<sup>-1</sup>. These impressive electrochemical properties indicate that ZGAs have substantial prospect as superior-performance anode materials for LIBs.<sup>146</sup>

In the study described in ref. 198, cube-like, interconnected, amorphous ZnSnO<sub>3</sub>@TiO<sub>2</sub> nanocomposites were fabricated using a simple co-precipitation followed by TiO<sub>2</sub> coating

procedure. The incorporation of an interconnected TiO<sub>2</sub> wrapping and numerous mesopores in the composite structure served to alleviate volume expansion stress, resulting in significantly improved electrochemical stability and performance. The hierarchical heterostructure of the ZnSnO<sub>3</sub>@TiO<sub>2</sub> nanocomposite exhibited impressive electrochemical characteristics. It demonstrated a high initial discharge capacity of 1590 mA h g<sup>-1</sup> at a current density of 100 mA g<sup>-1</sup> and a capacity retention of 780 mA h g<sup>-1</sup> after 200 cycles, surpassing that of pure ZnSnO<sub>3</sub> nanomaterial anodes. The benefits of the hierarchical nanocomposite morphology, together with the synergistic activity of ZnSnO<sub>3</sub> and TiO<sub>2</sub> contributed to the improved cycle life of the ZnSnO<sub>3</sub>@TiO<sub>2</sub> anode. Several factors contribute to its improved performance, as follows: (1) the hierarchical porous 3D morphology reduced the transportation length of lithium ions and electrolyte molecules, increased the contact area between the electrolyte and active material, and enhanced the penetration capability of the electrolyte molecules. (2) The hierarchical porous morphology helped mitigate the significant volume enlargement that occurs during repetitive charge-discharge cycles. (3) The TiO<sub>2</sub> coating layer effectively alleviated the stress induced by volume changes during the Li<sup>+</sup> insertion-extraction conversion reaction. (4) The synergistic effect of the high capacity of ZnSnO<sub>3</sub> and superior structural stability of TiO<sub>2</sub> contributed to the superior electrochemical performance of the ZnSnO<sub>3</sub>@TiO<sub>2</sub> composite. Overall, the hierarchical ZnSnO<sub>3</sub>@TiO<sub>2</sub> nanocomposite demonstrated promising electrochemical properties, making it a potential candidate for high-performance anode materials in LIBs.<sup>198</sup>

A sustainable cubic ZnSnO<sub>3</sub>/carbon composite with a high surface area of 896.39 m<sup>2</sup> g<sup>-1</sup> was successfully synthesized using a low-cost combustion process.<sup>199</sup> The electrochemical

Table 6 Anode materials for lithium-ion batteries based on ZnSnO<sub>3</sub> and their performance

| Materials                                         | Initial capacity<br>(mA h g <sup>-1</sup> ) | Reversible capacity<br>(mA h g <sup>-1</sup> ) | Cycle number<br>(n) | Current density<br>(mA g <sup>-1</sup> ) | Capacity retention<br>(%) | Ref. |
|---------------------------------------------------|---------------------------------------------|------------------------------------------------|---------------------|------------------------------------------|---------------------------|------|
| Sandwich-like ZnSnO <sub>3</sub> @C               | 1573                                        | 1107                                           | 200                 | 100                                      | 70                        | 194  |
| Sheet-like ZnSnO <sub>3</sub>                     | 1321                                        | 872                                            | 100                 | 100                                      | 66                        | 194  |
| ZnO@ZnSnO <sub>3</sub> core-shell                 | 1275                                        | 1073                                           | 110                 | 200                                      | 84                        | 197  |
| Amorphous ZnSnO <sub>3</sub> @C                   | 1633                                        | 1060                                           | 100                 | 200                                      | 65                        | 196  |
| Crystalline ZnSnO <sub>3</sub> @C                 | 1350                                        | 791                                            | 100                 | 200                                      | 59                        | 196  |
| Cube-like ZnSnO <sub>3</sub> @TiO <sub>2</sub>    | 1590                                        | 780                                            | 200                 | 100                                      | 49                        | 198  |
| ZnSnO <sub>3</sub> /rGO                           | 1987.5                                      | 745                                            | 100                 | 100                                      | 38                        | 146  |
| ZnSnO <sub>3</sub> /rGO                           | 2013                                        | 718                                            | 100                 | 100                                      | 36                        | 145  |
| ZnSnO <sub>3</sub> /rGO                           | 1691                                        | 713                                            | 100                 | 100                                      | 42                        | 200  |
| S@Ni/NiO/ZnSnO <sub>3</sub>                       | 1070                                        | 493                                            | 400                 | 335                                      | 46                        | 201  |
| S@ZnSnO <sub>3</sub>                              | 781                                         | 312                                            | 400                 | 335                                      | 40                        | 201  |
| ZnSnO <sub>3</sub> /C nanofibers                  | 1412                                        | 586                                            | 100                 | 100                                      | 42                        | 73   |
| ZnSnO <sub>3</sub> /C                             | 1469                                        | 695                                            | 200                 | 100                                      | 57                        | 189  |
| ZnSnO <sub>3</sub> @CNF                           | 1183                                        | 582.6                                          | 100                 | 100                                      | 49                        | 188  |
| ZnSnO <sub>3</sub> nanocubes                      | 1470                                        | 590                                            | 50                  | 200                                      | 40                        | 101  |
| Double-shell-ZnSnO <sub>3</sub> hollow microcubes | 2134                                        | 741                                            | 50                  | 100                                      | 35                        | 195  |
| Yolk-shell ZnSnO <sub>3</sub>                     | 1570                                        | 628                                            | 50                  | 100                                      | 40                        | 195  |
| ZnSnO <sub>3</sub> -C hollow microcubes           | 2255                                        | 703                                            | 50                  | 100                                      | 31                        | 202  |
| ZnSnO <sub>3</sub> amorphous hollow microcubes    | 1921                                        | 400                                            | 50                  | 100                                      | 21                        | 202  |
| Sheet-like ZnSnO <sub>3</sub>                     | 1390                                        | 625                                            | 50                  | 100                                      | 45                        | 71   |
| ZnSnO <sub>3</sub> @C/rGO                         | 1984                                        | 1040                                           | 45                  | 100                                      | 52                        | 203  |
| Ag/ZnSnO <sub>3</sub>                             | 874.5                                       | 464.5                                          | 45                  | 300                                      | 53                        | 204  |



performance of supercapacitors using the  $\text{ZnSnO}_3/\text{C}$  nanocomposite as electrode materials was demonstrated, revealing a good rate capability and exceptional cycling stability. The presence of  $\text{ZnSnO}_3$  in the composite facilitated rapid ion transfer, resulting in a capacitance retention of approximately 80% after 3000 charge–discharge cycles at different current densities in 6 M KOH electrolyte. These results indicate that the  $\text{ZnSnO}_3$ /carbon composite shows promise as a lightweight electrode material for supercapacitors due to its quick charge–discharge characteristics and strong endurance.

Furthermore, it is anticipated that unique morphology of anode materials, such as porous hollow nanostructures<sup>198</sup> and hybrid carbon-based nanocomposites,<sup>199</sup> can be a viable approach for enhancing lithium storage. However, it is necessary to improve the existing system for evaluating anode materials and assessing their charge–discharge capacity (Table 6).

#### 4.4. Nanodevices

Materials with piezoelectric properties have the remarkable ability to transform mechanical energy into electrical energy. In 2006, Wang<sup>206</sup> pioneered the concept of piezoelectric nanogenerators (PENGs), and since then, there has been an explosion of innovation in the design of nanostructured materials to serve as PENGs for harnessing diverse forms of energy from their surroundings.<sup>121,206,207</sup> This surge in research has led to the development of various piezoelectric materials, including inorganic materials, to further advance this field.<sup>205,208–210</sup>

In the study conducted by Wang *et al.*, a lead-free nanogenerator was developed using a  $\text{ZnSnO}_3$  microbelts, which demonstrated an impressive performance.<sup>36</sup> When subjected to

compressive and releasing strain ranging from 0.8% to 1.0%, the  $\text{ZnSnO}_3$  microbelts, anchored at their ends on a flexible polystyrene substrate, produced an output power of approximately 3 nW. The resulting output voltage and current were measured to be 100 mV and 30 nA, respectively. These values correspond to an energy conversion efficiency in the range of 4.2% to 6.6% based on the applied strain of 0.8% to 1%. These findings highlight the significant advancements in lead-free  $\text{ZnSnO}_3$  nanogenerators, showcasing their potential for high output power generation and efficient energy conversion through piezoelectric effects.

Datta and colleagues made a significant contribution by demonstrating the remarkable ferroelectric properties of  $\text{ZnSnO}_3$  nanowire arrays.<sup>211</sup> They discovered that these arrays exhibited an exceptional remanent polarization value of approximately  $30 \mu\text{C cm}^{-2}$ . This groundbreaking finding provides valuable insights for the future development of  $\text{ZnSnO}_3$ -based ferroelectric devices, enabling a more coordinated and targeted design approach.

Piezoelectric nanogenerators (PENGs) are highly valued for their ability to convert ambient mechanical stimuli into electrical energy, making them ideal for self-powered sensors. In ref. 212, the researchers employed a low-temperature solution-based approach to synthesize aluminium-doped zinc stannate ( $\text{ZnSnO}_3$ ) PENGs, which exhibited high electrical outputs in response to external forces. The improved output was ascribed to several factors, including the modified crystalline structure of  $\text{ZnSnO}_3$ , the incorporation of aluminium cations with specific modifications, and the reduced size of the  $\text{ZnSnO}_3$  nanocube particles. Through their experiments, the researchers

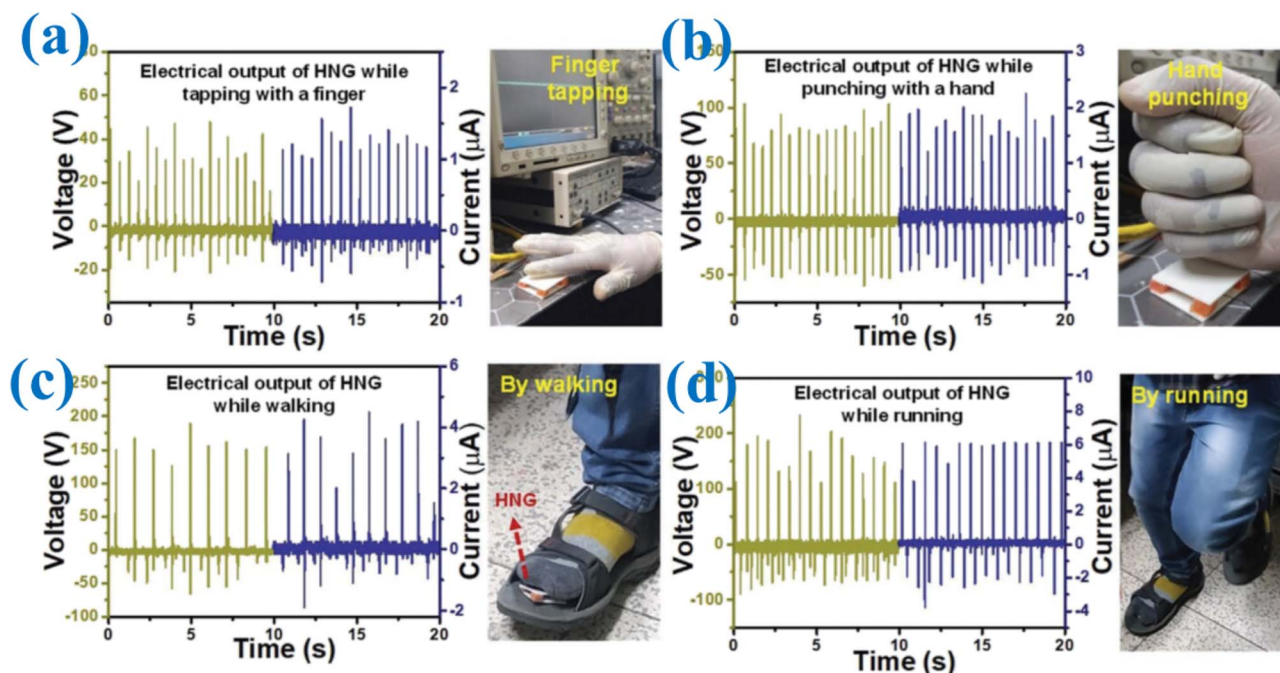


Fig. 21 HNG device showing promising potential for real-time and practical applications in harvesting mechanical energy from everyday human activities. Visual illustrations and the associated electrical data showing the performance and output of the HNG device for (a) finger tapping, (b) hand punching, (c) walking, and (d) running.<sup>213</sup>

determined that the optimal concentration of aluminium dopant was 2 wt%. The resulting Al-doped  $\text{ZnSnO}_3$  PENG generated impressive electrical outputs of 110 V, 13  $\mu\text{A}$ , and  $1200 \text{ mW m}^{-2}$  with just 120 beats per min (2 Hz) finger tapping as the input. Remarkably, the PENG charged a  $2.2 \mu\text{F}$  capacitor to 3 V in a mere 8 s with the simple act of finger tapping. Leveraging the high output-to-input ratio, the researchers integrated the PENG with a helmet, enabling it to function as a motion sensor. By detecting the mechanical movements of the wearer, the device converted each footstep into a wireless signal, showcasing its potential applications in capturing human motion data.

The improved electrical output of nanogenerators is attributed to the enhanced dielectric features, enhanced electric dipole moment, and piezoelectric coefficient exhibited by piezoelectric materials. To harness the synergistic effects of  $\text{ZnSnO}_3$  and polydimethylsiloxane (PDMS) polymer, Yu and colleagues synthesized a composite polymer matrix (CPM) by embedding  $\text{ZnSnO}_3$  nanocrystals (NCs) in triboelectric PDMS.<sup>213</sup> The resulting hybrid nanogenerator (HNG) demonstrated an impressive performance, with open-circuit voltage ( $V_{\text{OC}}$ ), short-circuit current ( $I_{\text{SC}}$ ), power density, and charge density values of approximately  $\sim 410 \text{ V}$ , 13  $\mu\text{A}$ ,  $2 \text{ W m}^{-2}$ , and  $95 \text{ C m}^{-2}$ , respectively, for the 0.15 g reduced graphene oxide (rGO)- $\text{ZnSnO}_3$ /PDMS CPM-based HNG. The electrical energy produced by harvesting mechanical energy using the HNG was investigated during various human body actions, including finger tapping, hand punching, walking, and running, as depicted in Fig. 21a–d. The obtained electrical output, as represented by the  $V_{\text{OC}}/I_{\text{SC}}$  curves of the HNG, was approximately  $\sim 40 \text{ V}/1.5 \mu\text{A}$ ,  $\sim 80 \text{ V}/2 \mu\text{A}$ ,  $\sim 160 \text{ V}/4 \mu\text{A}$ , and  $\sim 190 \text{ V}/6 \mu\text{A}$  for tapping with a finger,

punching with a hand, walking, and running, respectively. These results demonstrate the remarkable potential of the HNG in converting mechanical energy from human actions into electrical energy.

Furthermore, the integration of single-crystalline piezoelectric perovskite  $\text{ZnSnO}_3$  nanocubes with PDMS for the construction of hybrid piezoelectric nanogenerators (HP-NGs) is a significant development. Lee *et al.* successfully demonstrated a highly stable and powerful HP-NG without the need for electrical poling treatment.<sup>214</sup> The HP-NGs incorporating  $\text{ZnSnO}_3$  nanocubes showed remarkable mechanical durability, robustness, and power-generation capabilities. To evaluate their piezoelectric power generation and mechanical resilience, a direct impact approach was employed. In one experiment, a large-scale HP-NG was placed on a road and subjected to vertical compressive force exerted by a heavy motor vehicle. Fig. 22 captures a snapshot of this experimental motion.

The observed the output voltage and current from the  $10 \text{ cm} \times 10 \text{ cm}$  HP-NG are depicted in Fig. 22c and d, demonstrating a notable recorded output voltage of approximately  $\sim 20 \text{ V}$  and output current density of about  $1 \mu\text{A cm}^{-2}$ .

The utilization of  $\text{ZnSnO}_3$  triangular belts as lead-free nanogenerators has yielded groundbreaking results [Z]. By applying a mechanical deformation of 0.1%, these nanogenerators could generate a current of  $0.13 \mu\text{A}$  and voltage of  $5.3 \text{ V}$ , marking the first verification of high output power from lead-free  $\text{ZnSnO}_3$  nanogenerators.<sup>37</sup>

Fig. 23 presents a schematic diagram illustrating the working mechanism of generating power from a piezoelectric nanogenerator (PENG). Research has confirmed that when

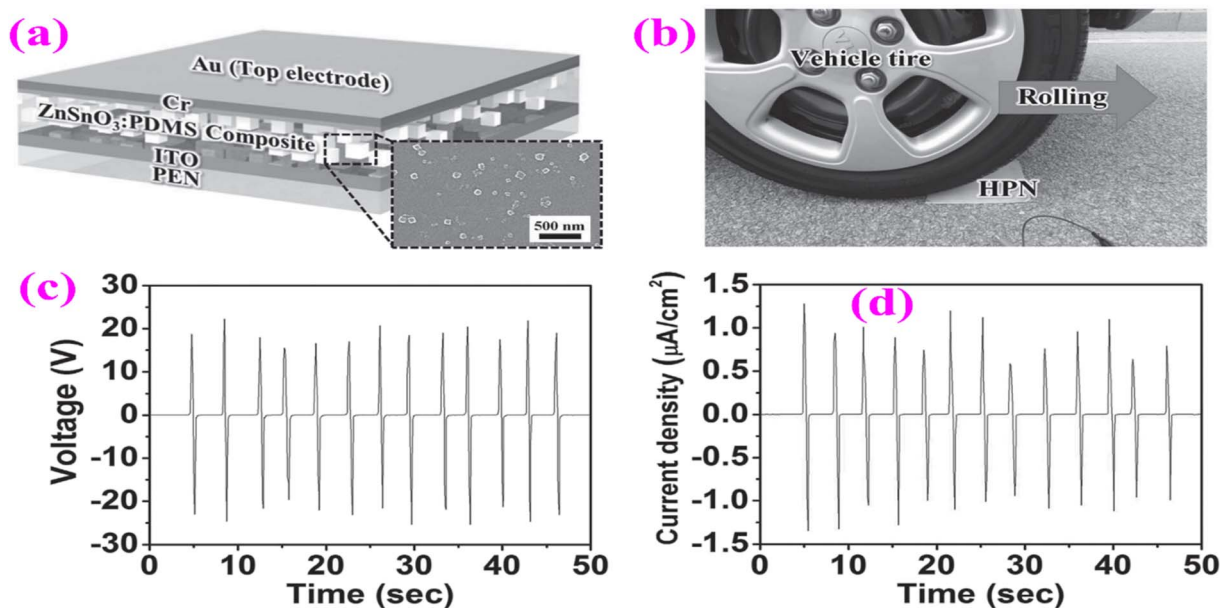


Fig. 22 Structure and power generation by HP-NG when a tire is rolled. (a) Illustration of the structure of the  $\text{ZnSnO}_3$ :PDMS nanocomposite-based flexible HP-NG as seen by the cross-sectional FESEM image. (b) Experimental setup for power production under the rolling of a vehicle tire, in which  $\text{ZnSnO}_3$ :PDMS HP-NG was attached to the road. (c) Output voltage and (d) current density from HP-NG induced by the loading and unloading of a vehicle tire.<sup>214</sup>



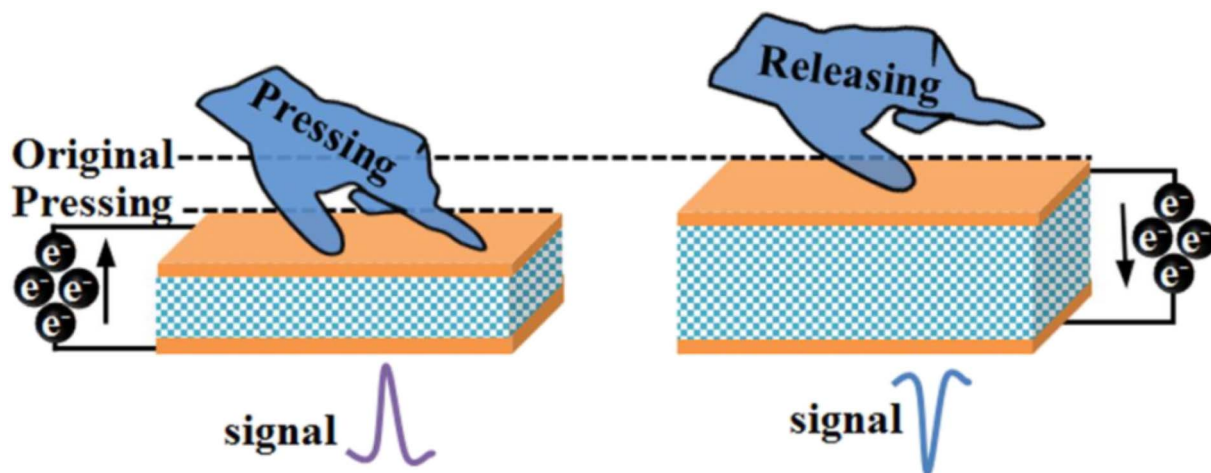


Fig. 23 Working mechanism of generating power from a piezoelectric nanogenerator.<sup>215</sup>

subjected to longitudinal compressive stress, the electric dipoles in the crystal align effectively in a single direction due to the phenomenon known as the stress-induced poling effect.<sup>215</sup> When a tapping force is applied to the device, this pressure causes the separation of positive and negative charges in the ZnSnO<sub>3</sub> nanorods, creating an electric dipole.<sup>216</sup> After the cancellation of the interior charges with opposite signs, excess charges remain at both surfaces, leading to a net polarization effect, as depicted in Fig. 7b. This resulting polarization induces opposite charges at the corresponding electrodes, initiating the flow of electrons through the external circuit, thereby generating voltage and current. This mechanism generates a piezoelectric potential across the surface, allowing free electrons to move from the top and bottom electrodes through the external circuit. To balance this generated piezoelectric potential, the electrons accumulate at the interfaces between the nanocubes and electrodes. Subsequently, when the nanocubes decompress after removing the applied load or object, the accumulated electrons return through the external load, and the piezoelectric potential is neutralized.<sup>215,216</sup>

On pressing it will give a positive peak, whereas on releasing the pressure, it produces a negative peak as the electrons move in the opposite direction. Higher and smaller peaks were also observed when compression occurred more rapidly than decompression and the magnitude of the both output voltage and current strongly depends on the compression frequency and applied pressure.<sup>210,215,216</sup>

## 5. Conclusion and perspectives

Nanostructured materials based on ZnSnO<sub>3</sub> possess a multitude of features. These include optical properties, a high isoelectric point, large specific area, non-toxic nature, mechanical characteristics, bending strength, and flexibility. Each of these properties is significant and unique, contributing to the functionality and potential applications of ZnSnO<sub>3</sub>-based nanostructures. Researchers have dedicated their efforts to exploring the fundamental aspects of developing, modifying, and

utilizing ZnSnO<sub>3</sub>-based nanostructures in various energy and biological applications due to their intriguing characteristics. In this study, we focused on ZnSnO<sub>3</sub>-based nanostructures and provided a comprehensive overview, with particular emphasis on their applications in Li-ion batteries, photocatalysis, biosensors, and the biological domain. Our review primarily addresses the advancements made in ZnSnO<sub>3</sub>-based materials to achieve higher performances across diverse applications.

ZnSnO<sub>3</sub>-based materials exhibit photocatalytic properties, enabling the decomposition of dye molecules upon exposure to visible light. The efficiency of catalysts in photodegradation and their long-term stability are influenced by factors such as structural morphology, crystal orientation, particle size, and oxygen defects. To enhance the photocatalytic activity of ZnSnO<sub>3</sub>, several strategies have been employed, including the fabrication of ZnSnO<sub>3</sub> nanostructured arrays, incorporating additional components such as noble metals, introducing dopants, and integrating carbon materials. These approaches aim to improve the overall photocatalytic performance of ZnSnO<sub>3</sub>-based materials. However, there are still challenges and opportunities in the field of ZnSnO<sub>3</sub> photocatalysis. Despite extensive research in controlled environments, key factors such as surface excess charges, surface additives, and the states of adsorbates in the photocatalytic process are not well-understood. Additionally, there is no consensus on the nature of photogenerated charge carriers and their relationship with charge/energy transfer and bond breaking/forming due to a lack of detailed knowledge about their interactions with adsorbates. Thus, to address these issues, it is crucial to complement traditional ensemble-averaged experimental techniques with time-resolved methods. These experiments aim to uncover the dynamic aspects of ZnSnO<sub>3</sub> photocatalysis, leading to a deeper understanding and the potential development of new dynamic-based photocatalysis models.

Researchers have shown that ZnSnO<sub>3</sub>-based nanostructured materials possess the capability to bind gaseous molecules. In sensing applications, the utilization of nanostructures aims to increase the surface-to-volume ratio, enhancing their reactivity



and facilitating their interaction with analytes by facilitating their translocation across cell membranes. To satisfy to bio-sensing and environmental requirements, scientists have developed electrochemical sensors utilizing ZnSnO<sub>3</sub> nanostructures and composite materials. These advancements open possibilities for improved gas sensing and environmental monitoring applications. Due to the benefits mentioned earlier, we emphasize the synthesis strategy to enhance the efficiency of ZnSnO<sub>3</sub> and optimize its operating temperature. It is advisable to investigate techniques that can produce ZnSnO<sub>3</sub> structures characterized by significant porosity and substantial specific surface areas. Given that surface resistance plays a crucial role in gas sensing, the creation of highly porous structures is advantageous, offering additional reactive sites for the absorption and interaction of the desired gas molecules, and thereby promoting the formation of an electron depletion region. Despite the extensive use of ZnSnO<sub>3</sub> nanomaterials and their composites for gas sensing, their selectivity and high operating temperature remain questionable. Thus, to enhance selectivity for specific gases, it is recommended to conduct dynamic tests at varying temperatures. Different temperatures influence the adsorption and desorption processes of distinct gas molecules on the material surface, resulting in temperature-dependent interactions between the sensing material and specific gas molecules. This dynamic testing approach enables the confirmation of the selectivity of a sensor for a particular gas at a specific temperature. Consequently, in practical applications, a single sensor can detect multiple gases by adjusting the working temperature, enabling the accurate identification of the components in mixed gases and significantly expanding the detection capabilities of the sensor.

ZnSnO<sub>3</sub>-based nanostructures have been found to exhibit promising energy storage characteristics, indicating their potential use as electrode materials in supercapacitors and lithium-ion batteries. ZnSnO<sub>3</sub> is widely recognized as a viable active material for batteries, demonstrating a typical reversible energy of 1107 mA h g<sup>-1</sup> after 200 cycles. These findings highlight the suitability of ZnSnO<sub>3</sub> for energy storage applications and suggest its potential for enhancing the performance of supercapacitors and LIBs. Although various related structures have been created and considerable progress has been made in energy storage, there are still many challenges that need further exploration. First, cleverly designed structures optimize certain properties of LIBs, but simultaneously these anodes present other issues. For example, porous or hollow structures can accommodate the volume changes but reduce the volumetric capacity of the whole cell. For the prevalent application of portable, light, small electric devices, volumetric capacity and specific capacity are highly valued.

The comprehensive assessment of anode materials and structures is necessary, aiming to produce electrode architecture designs endowed with better overall performances. Although composites of multi-material blending and combination of different structures are beneficial, more theoretical research including modelling and computing should be applied to maximize the features leading to improved performances,

while also minimizing the shortcomings among the different materials and structures.

The exceptional properties of nanostructured ZnSnO<sub>3</sub>-based materials in energy storage, biosensor, biological, and PENG applications highlight their potential for further scientific advancements in interdisciplinary research. There is significant scope for enhancing the performance of ZnSnO<sub>3</sub>-based materials in the energy, environmental, and biology fields. Accordingly, further investigations into the tunable characteristics of these materials are anticipated, aiming to unlock their full potential and contribute to advancements in these diverse fields.

## Author contributions

Moksodur Rahman: conceptualization, formal analysis, investigation, methodology, visualization, writing – original draft; Muhammad Shahriar Bashar: supervision, resources, validation, writing – review & editing; Md. Lutfor Rahman: validation, visualization, writing – review & editing; Faisal Islam Chowdhury: conceptualization, supervision, project administration, writing – review & editing.

## Conflicts of interest

There are no conflicts to declare.

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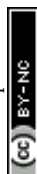
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