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Highly active ZIF-8@CNT composite catalysts as cathode materials for anion exchange membrane fuel cells†

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Developing non-precious metal-based inexpensive and highly active electrocatalysts for the oxygen reduction reaction (ORR) in alkaline media is important for fuel cell applications. Herein, we report a simple and effective synthesis of transition-metal-doped zeolitic imidazolate framework-8 (ZIF-8) and carbon nanotube (CNT) composite catalysts (ZIF-8@CNT) prepared *via* high-temperature pyrolysis at 900 °C. The catalysts were characterized using different physicochemical techniques and employed as cathode materials in anion exchange membrane fuel cells (AEMFC). The prepared metal-free (ZNT-900), single-metal-doped (Fe-ZNT-900, Co-ZNT-900) and binary-metal-doped (Fe₁Co₁-ZNT-900, Fe₁Co₂-ZNT-900) catalysts had a sufficient amount of N-doping with the presence of FeCo moieties in the carbon skeleton of the latter two materials. N₂ adsorption-desorption isotherms showed that all the prepared catalysts possess a sufficient Brunauer-Emmett-Teller surface area with more micropores present in ZNT-900, while a combined micro-mesoporous structure was obtained for transition-metal-doped catalysts. Binary-metal-doped catalysts showed the highest number of ORR-active sites (pyridinic-N, pyrrolic-N, graphitic-N, M-N_x) and exhibited a half-wave potential ($E_{1/2}$) of 0.846 and 0.847 V vs. RHE for Fe₁Co₁-ZNT-900 and Fe₁Co₂-ZNT-900, respectively, which surpassed that of the commercial Pt/C catalyst ($E_{1/2}$ = 0.834 V). In H₂-O₂ AEMFCs, the Fe₁Co₂-ZNT-900 catalyst delivered a maximum power density (P_{\max}) of 0.171 W cm⁻² and current density at 0.5 V ($j_{0.5}$) of 0.326 A cm⁻², which is very close to that of the Pt/C catalyst (P_{\max} = 0.215 W cm⁻² and $j_{0.5}$ = 0.359 A cm⁻²). The prepared ZIF-8@CNT catalysts showed remarkable electrocatalytic ORR activity in 0.1 M KOH solution and fuel cell performance comparable to that of the benchmark Pt/C catalyst.

Keywords: Rotating disk electrode; Anion exchange membrane fuel cell; Zeolitic imidazolate framework; Non-precious metal catalyst; Oxygen reduction reaction.

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1. Introduction

To mitigate global warming, current research and development have received a lot of attention in the field of renewable energy generation.^{1–5} Among promising electrochemical energy conversion devices are anion exchange membrane fuel cells (AEMFC) that use O₂/air as an oxidant at the cathode to produce electric power and some heat.⁶ However, the sluggish kinetics of cathodic oxygen reduction

reaction (ORR) and insufficient durability are significant disadvantages in AEMFC practical applications.^{7,8} Currently, platinum (Pt) represents the pinnacle of electrocatalytic ORR;⁹ however, Pt-based catalysts are restricted by its scarcity, vulnerability to carbon monoxide (CO) poisoning, *etc.*¹⁰ Research on non-platinum catalyst materials is crucial for the widespread development of AEMFCs.^{11–15} To date, numerous approaches have been used to make progress in this direction to increase the density¹⁵ of ORR-active sites *via* introducing multiple active catalyst centres.¹⁶ Various transition-metal-based electrocatalysts have been explored as cost-efficient materials for catalysing the ORR in alkaline medium; nonetheless, it is challenging to identify a single-metal electrocatalyst as a counterpart to Pt group metal catalysts.^{17–19} In this regard, transition metal and nitrogen co-doped carbon (M-N-C)-based catalyst materials have attracted great

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attention due to their excellent ORR activity in alkaline media, which sometimes surpasses that of the Pt/C catalyst.²⁰ The M–N–C catalysts are generally prepared by pyrolysing the carbon-, nitrogen- and transition-metal-containing precursors.²¹ Typically, M–N₄ species are believed to be the most active ORR centres.²² According to many reports, the better performance of single-atom-based Fe–N–C catalysts usually coming from the well-established Fe–N_x or porphyrin-like Fe–N₄ active sites obtained after high-temperature pyrolysis.^{22–24} Bimetallic catalysts with controlled electronic structures and enhanced synergistic effects also show extremely strong potential for ORR application. Bimetallic electrocatalysts are based on the structural modification between two metal atoms due to which the resulting chemical structure is expected to have enhanced electronic properties favourable for increased electrocatalytic activity.^{25,26} Reports suggest that the bimetallic sites are believed to facilitate the O–O bond cleavage *via* a bridge *cis* adsorption of ORR intermediates and also improve the durability of M–N–C sites for the ORR.²⁷ Therefore, the construction of hybrid electrode materials, especially those made from Co and Fe metal sources (including Fe/Co bimetallic oxides and alloys), has become the main focus nowadays.^{28–30} These bimetallic catalysts can be synthesised through thermal conversion of bimetallic sources with various Fe/Co ratios. In recent years, to reduce the cost, non-precious-metal catalysts derived from Co, Fe, Mn, *etc.* are usually used as anode as well as cathode catalysts for various applications.³¹ Additionally, metal–organic frameworks (MOFs) have significant advantages in catalyst tuning and structure optimization and their derivatives usually have highly dispersed accessible active sites, improved charge transfer networks, and high porosity for efficient mass transfer of reactants.^{31–33} Zeolitic imidazolate framework-8 (ZIF-8), as a branch of MOFs, has gained much attention as a template for preparing carbon-based catalysts due to its hierarchical porosity, high nitrogen content, excellent thermal and electrochemical stability, and volatile characteristics of Zn at higher temperatures.^{34,35} This further enhances doping of nitrogen in the carbon skeleton, which changes the charge distribution and helps O₂ adsorption on the active sites.^{36,37}

To date, a vast array of monometallic and bimetallic Fe- and Co-doped nitrogen-containing nanocarbon catalysts have been researched in AEMFCs. For instance, our group reported the potential role of Fe and Co doping on graphene/CNT composite catalysts in enhancing the power output of AEMFCs.³⁸ The prepared composite carbon catalysts showed an excellent ORR activity with a half-wave potential ($E_{1/2}$) reaching 0.81 V *vs.* RHE, and an AEMFC power output of 0.64 W cm^{−2} was achieved for the bimetallic Fe- and Co-doped catalysts. Another study was reported with Fe- and Co doped mesoporous carbon catalysts applied as cathode materials in AEMFCs. The bimetallic catalysts reached an $E_{1/2}$ value of 0.86 V and a maximum power density of 0.92 W cm^{−2} in AEMFCs.³⁹ Fe and Co metal anchored ZIF-8-based carbon catalysts are not extensively studied in AEMFCs. The anion exchange membrane (AEM) plays a crucial role in overall

AEMFC performance and many studies so far have been performed with the Aemion+ AEM and ionomer, which offers a lower mass transport resistance and water management throughout the AEMFC.⁴⁰ To our knowledge, commercial Aemion+® 15 µm AEM has not been tested in AEMFC and this study is the first of its kind to show the potential role of the Aemion+® 15 µm membrane in AEMFC performance. To our knowledge, a hybrid composite of ZIF-8 and CNT derived Fe–Co catalysts and their demonstration in the AEMFC is reported for the first time.

Herein, we report highly dispersed monometallic (Fe or Co) and bimetallic FeCo catalysts prepared by the composite of CNTs and N-rich ZIF-8. The resulting catalyst materials exhibit superior electron transfer properties and high electrocatalytic activity towards the ORR. Amongst the ZIF-8@CNT catalysts, FeCo bimetallic/N-doped CNT catalysts reached the highest $E_{1/2}$ value and surpassed the performance of commercial Pt/C (20 wt% Pt) catalysts and most of the reported Fe/Co-based MOF catalysts in 0.1 M KOH solution. The performance of bimetallic FeCo-based catalysts was further investigated at the AEMFC cathode, and for Fe₁Co₂-ZNT-900 (Fe/Co at a ratio of 1:2, respectively) the peak power density of 0.171 W cm^{−2} was obtained, which is close to that of Pt/C using the Aemion+® 15 µm AEM for AEMFC. The adopted synthesis approach to design efficient FeCo bimetallic cathode catalysts for AEMFC applications would drive towards sustainable AEMFC technologies.

2. Results and discussion

2.1. Physicochemical characterisation of ZIF-8@CNT catalysts

Fig. 1a shows the crystallographic features of as-prepared ZIF-8@CNT catalysts revealed by X-ray diffraction (XRD) analysis. The appearance of broad diffraction peaks at 25.6° and 43.08° represents the (002) and (100) graphitic planes (PDF-04-020-4354), proving that ZIF-8@CNT catalysts have been successfully pyrolysed and carbonised.⁴¹ At high-temperature annealing (900 °C), amorphous carbon disappears and graphitised carbon peaks are dominant.⁴² The characteristic peaks located at 30.13° (220), 35.55° (311), 57.28° (511) and 62.94° (440) suggest the presence of magnetic Fe₃O₄ nanoparticles (PDF-00-039-1346).⁴³ Also, a peak around 44.67° can be assigned to the α -Fe body-centred (110) plane (PDF-04-007-9753).⁴⁴ Two characteristic peaks at 44.89° and 65.4° can be attributed to FeCo alloys residing in the (111) and (200) lattice planes (PDF-00-049-1568).⁴⁵

The defects in the as-synthesised ZIF-8@CNT catalysts were analysed by Raman spectroscopy (Fig. 1b). The D band displayed at ~1352 cm^{−1} shows the defects and disorders in the carbon lattice, and the G band at ~1584 cm^{−1} reflects the E_{2g} vibrations in the graphite-type lattice (sp²-hybridised carbon).^{46,47} For MWCNTs, the G band position is found to be higher than that of other carbon-based materials, which is reported mainly due to large compressive stress affecting the C=C bonds present in highly defective nanotube walls. Further deconvolution of the D and G bands provides





Fig. 1 (a) XRD patterns, (b) Raman spectra, (c) N₂ adsorption-desorption isotherms and (d) pore size distribution for various transition-metal-doped and metal-free ZIF-8@CNT-derived catalysts.

additional information regarding the amorphous carbon nature. For instance, Daniel *et al.* presented polyurethane-derived Fe-N-C electrocatalysts for the ORR.⁴⁸ The Raman spectra showed two broader peaks centred at ~1350 and ~1595 cm⁻¹ indicating D and G bands, respectively, and were further deconvoluted to six Gaussian bands named I, D4, D1, D3, G and D2. However, deconvoluting Raman spectra in the case of MWCNTs is more complicated and is not often studied.⁴⁹ It has been reported that the complex structure of CNTs with different scattering events happened due to weak out-of-plane modes, weak wall-to-wall interactions, and curvature of walls which makes their deconvolution quite challenging.⁵⁰ As an example, all spectra were fitted following the five Voigt-shaped band model, where the G band (~1585 cm⁻¹) corresponds to the stretching vibrations of the sp² carbon atoms in the ideal graphitic lattice; D1 (~1350 cm⁻¹, Lorentzian) to defect-activated breathing mode of aromatic rings; D2 (~1610 cm⁻¹, Gaussian) to disordered graphene layers at the surface of a graphitic crystal; D3 (~1490 cm⁻¹, Gaussian) to amorphous carbon and D4 (~1210 cm⁻¹, Gaussian) to disordered graphitic lattice. For instance, the deconvoluted Raman spectrum for Fe₁Co₂-ZNT-900 is presented in Fig. S2.† It should be considered that the commonly used I_D/I_G Raman band peak intensity ratio is rather appropriate to quantify the structural imperfection for

the graphene samples. However, for highly disordered carbon materials, the Raman band commonly labelled as G is in fact a superposition of the G and D2 bands. This band overlap renders the I_D/I_G ratio an unreliable measure for the structural disorder of such materials.⁵¹ Therefore, to interpret the degree of structural disordering in ZIF@CNT catalysts, the parameter W (full-width at half-maximum (FWHM) of the D1 and G bands), I_{D1}/I_G (the ratio of the areas under the bands) and $R2 = D1/(G + D1 + D2)$ (the ratio of the areas of indicated bands) were calculated as shown in Table S2.† The amorphous carbon represented by band D3 was found similar for all the catalyst materials. In fact, the change in intensities between the D and G bands is mostly due to a variation of the widths of the D1 and G bands. From the data in Table S2.† it can be concluded that the structural disorder increases steadily in the sequence of Fe-ZNT-900, ZNT-900, Co-ZNT-900, Fe₁Co₁-ZNT-900, and Fe₁Co₂-ZNT-900 samples. It has been reported that pyrolysis promotes the graphitisation of the materials in the presence of small amounts of transition metals, which is evident from the higher graphitic peaks obtained for metal-doped catalysts.⁵²

The more defect-rich carbon structure obtained for bimetallic ZIF-8@CNT catalysts offers a higher number of active centres to promote the ORR kinetics.⁵³ The specific surface area and pore size distribution (PSD) of the ZIF-



Table 1 BET surface area (S_{BET}), volume of micropores (V_{micro}), total volume (V_{tot}), DFT surface area for micropores and mesopores calculated for ZIF-8@CNT catalysts

Catalysts	S_{BET} ($\text{m}^2 \text{g}^{-1}$)	V_{micro} ($\text{cm}^3 \text{g}^{-1}$)	V_{tot} ($\text{cm}^3 \text{g}^{-1}$)	S_{DFT} (mic.) ($\text{m}^2 \text{g}^{-1}$)	S_{DFT} (mes.) ($\text{m}^2 \text{g}^{-1}$)
ZNT-900	290	0.07	0.65	189	172
Fe-ZNT-900	334	0.07	0.66	145	170
Co-ZNT-900	297	0.07	0.73	141	175
Fe_1Co_1 -ZNT-900	247	0.05	0.61	106	147
Fe_1Co_2 -ZNT-900	312	0.06	0.66	109	181

8@CNT catalysts were investigated *via* N_2 adsorption-desorption isotherms and the Brunauer–Emmett–Teller (BET) theory, as shown in Fig. 1c and d, respectively. All the catalysts followed a combination of type I and III H3 hysteresis, which characteristically represents the presence of micro-mesoporous materials with a majority of mesopores according to IUPAC nomenclature.⁵⁴ The density functional theory (DFT) specific surface area (S_{DFT}) for the micropores (S_{DFT} mic.) and mesopores (S_{DFT} mes.) were estimated from N_2 isotherms using a quenched solid DFT (QSDFT) equilibrium model for slit-type pores (Table 1).

The ZNT-900 catalyst carried the highest volume of micropores and a lower BET surface area (S_{BET}). Since the ZNT-900 material contains an equal amount of CNTs along with ZIF-8, after annealing at 900 °C Zn evaporates, creating hollow ZIF crystals, and the ends of the CNTs start to close, which results in a decrease in its S_{BET} .⁵⁵ CNTs tend to form bundles due to van der Waals interactions and only the outer area, interstitial channels, and grooves are the available sites left for O_2 adsorption, which in turn decreases the overall surface area of the material. The DFT micropores/mesopores surface area ratio was close to 1.1 for ZNT-900, representing that both micropores and mesopores were equally present in the catalyst materials. The presence of Co did not affect S_{BET} because of the presence of more mesopores instead of micropores. A similar trend can be observed for the Fe_1Co_1 -ZNT-900 catalyst, in which bimetal doping obstructed the

micropores and increased the volume of mesopores, resulting in decreasing the S_{BET} even further to $247 \text{ m}^2 \text{g}^{-1}$. This can be attributed to the lower total pore volume ($0.61 \text{ cm}^3 \text{g}^{-1}$) and micropore volume ($0.05 \text{ cm}^3 \text{g}^{-1}$). Fe_1Co_2 -ZNT-900 showed an increase in the BET surface area up to $312 \text{ m}^2 \text{g}^{-1}$. In general, all the ZIF-8@CNT catalysts possess similar textural properties, especially the PSD, which shows the existence of both micro- and mesopores in the catalyst materials (Table 1). However, a small difference in the micropore/mesopore ratio can be observed for ZIF-8@CNT catalysts. These porous materials have a high potential for use as efficient electrocatalysts for the ORR and fuel cell applications.

The morphology and dopant distribution over the catalyst surface were studied in more detail using scanning electron microscopy (SEM) and scanning transmission electron microscopy (STEM) analysis. The surface morphologies of ZIF-8@CNT catalysts obtained from SEM images are shown in Fig. 2. All the catalysts showed similar morphologies with a tubular arrangement of CNTs covering the ZIF surface. The outer diameter of CNTs was in the range of 8–12 nm exhibiting a bamboo-like morphology with a smooth surface appearance. The incorporation of transition metals and CNTs in ZIF-8 displayed a visible difference, indicating the well-established interactions between ZIF-crystals and CNTs.

At lower magnifications (Fig. 2a–e), a hollow sponge-like structure of the ZNT-900 catalyst can be seen, whereas other

**Fig. 2** SEM images captured for (a and f) ZNT-900, (b and g) Fe-ZNT-900, (c and h) Co-ZNT-900, (d and i) Fe_1Co_1 -ZNT-900 and (e and j) Fe_1Co_2 -ZNT-900 catalysts at different resolutions.

transition-metal-doped ZIF-8@CNT catalysts contained a highly dense crystal structure of the catalysts, which was probably due to the presence of a micro-/mesopore combination. SEM-EDX mapping was performed to analyse the elemental composition of the as-prepared ZIF-8@CNT catalysts. The results revealed that all the catalysts consisted of similar amounts of carbon (Table 2) evenly distributed over the surface, which was as expected for the carbon-based catalysts (Fig. S3†). The surface metal concentrations disclosed by the SEM-EDX mapping and microwave plasma-atomic emission spectroscopy (MP-AES) analysis overlap with each other, showing a good incorporation of desired metal contents in the catalyst materials (Table 2). High-temperature annealing at 900 °C has been reported to effectively remove Zn traces from the material and the same was observed with our catalysts.^{56,57} SEM is confined to lower magnifications and does not provide an in-depth picture of the material. Therefore, STEM measurements were performed on all the catalyst materials at different resolutions. Bright-field (BF) STEM images for bimetallic ZIF-8@CNT catalysts (Fig. 3)

confirm the homogeneously wrapped ZIF 3D crystal structure with CNTs. The FeCo alloys can be clearly seen as dark dots in Fig. 3 in the images with different magnifications. It can be inferred that the dopants, *i.e.*, N, O, and transition metals, were well spread over the surface and metal nanoparticles of about 5 nm diameter were encapsulated in a ZIF-8@CNT porous carbon shell (Fig. S4†). Also, the high-angle annular dark field (HAADF)-STEM images captured at lower resolutions showed a bright spot corresponding to FeCo alloys (Fig. 4).

To further confirm the bonding configuration of the elements, X-ray photoelectron spectroscopy (XPS) analysis was performed. A comparative XPS survey spectrum recorded from 1000 to 200 eV for all the catalysts indicates the presence of different elements, *i.e.* C, N, O, Fe and Co, as shown in Fig. S5†. To identify the role of chemical entities, high-resolution XPS was carried out in the N 1s region (Fig. 5) and Fe 2p and Co 2p regions (Fig. S6†). A variety of nitrogen-bonded moieties were present in all the ZIF-8@CNT catalyst materials, namely, pyridinic-N (~398 eV), pyrrolic-N

Table 2 Surface elemental composition (wt%) of ZIF-8@CNT catalysts by SEM-EDX and MP-AES analysis

Catalyst	C	N	O	Fe (SEM-EDX)	Fe (MP-AES)	Co (SEM-EDX)	Co (MP-AES)
ZNT-900	90.42	6.45	1.96	0.00	—	—	—
Fe-ZNT-900	86.11	4.90	7.16	1.83	1.994	—	—
Co-ZNT-900	86.08	7.28	3.36	—	—	3.19	2.997
Fe ₁ Co ₁ -ZNT-900	89.16	3.17	4.46	2.70	3.561	2.86	3.016
Fe ₁ Co ₂ -ZNT-900	85.87	6.88	3.58	1.28	1.34	2.39	2.48



Fig. 3 BF-STEM images for (a–c) Fe₁Co₁-ZNT-900 and (d–f) Fe₁Co₂-ZNT-900 catalysts at different resolutions.



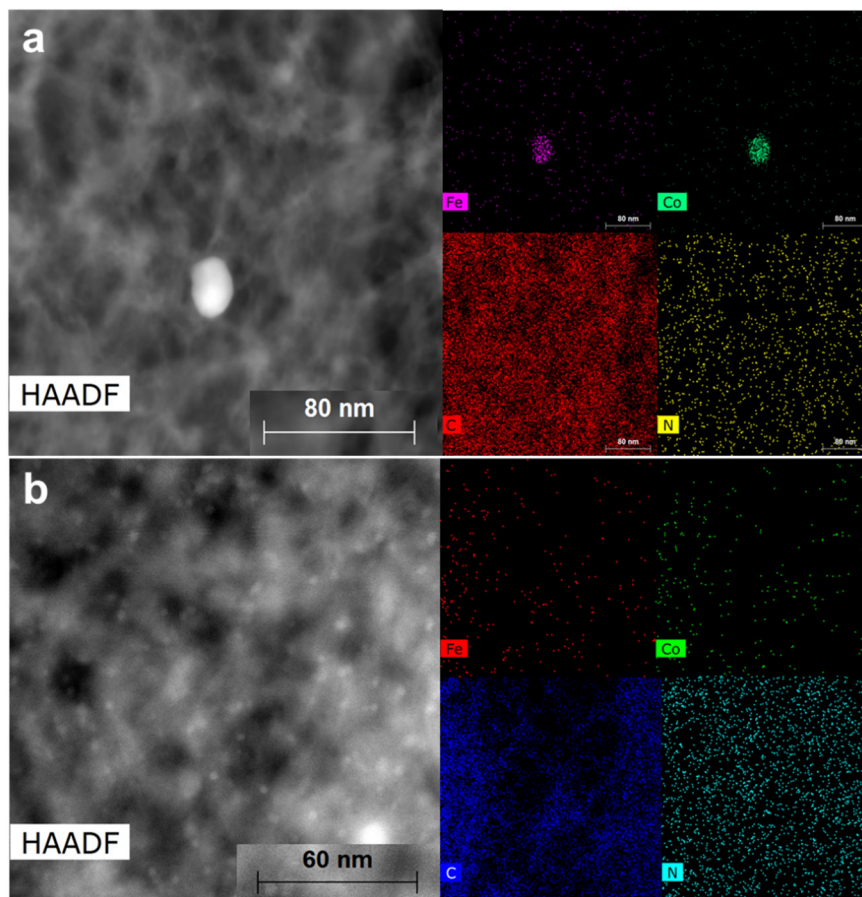


Fig. 4 HAADF-STEM images for (a) $\text{Fe}_1\text{Co}_1\text{-ZNT-900}$ and (b) $\text{Fe}_1\text{Co}_2\text{-ZNT-900}$ catalysts with corresponding EDX mapping.

(~ 400 eV), nitrogen-coordinated metal centres (M-N_x , ~ 399 eV), imines (~ 397.5 eV), graphitic-N (~ 401 eV), N-H (~ 403 eV), and finally oxidised nitrogen, N-O (~ 404 eV). The atomic ratios of all the elements are summarised in Table S3.† Several reports and theoretical molecular orbital diagram models suggest that the N atom transfers electrons to the π^* orbital of O_2 (ref. 58), and pyridinic-N, pyrrolic-N, M-N_x , and graphitic-N are the most encountered ORR-active centres.^{59–61} Therefore, the composition of these four N types are presented in Fig. 5f. The bimetallic ZIF-8@CNT catalyst ($\text{Fe}_1\text{-Co}_1\text{-ZNT-900}$ and $\text{Fe}_1\text{Co}_2\text{-ZNT-900}$) carries the highest amount of all four N species with a high amount of M-N_x centres, which promotes the electron transfer and, in turn, the ORR kinetics. The high-resolution XPS spectra in the Fe 2p region show doublets at ~ 710.8 and ~ 724.3 eV, reflecting the Fe 2p_{3/2} and Fe 2p_{1/2} spin-orbit coupling levels.⁶² The deconvoluted peaks of Fe 2p_{3/2} at binding energies of ~ 707.1 , ~ 710.8 and ~ 713.0 eV reflect different forms of Fe (Fe^0 , Fe^{2+} , Fe^{3+} , respectively). The peak at ~ 709.6 eV reflects the presence of Fe-N_x sites.^{63,64} It can be seen from Fig. S5† that the Fe-N_x intensity is highest in the $\text{Fe}_1\text{Co}_2\text{-ZNT-900}$ catalyst, reflecting its highly dense M-N_x centres. The two visible broader peaks at ~ 780.5 and ~ 796.0 eV indicate Co 2p_{3/2} and Co 2p_{1/2}, respectively. The deconvoluted peaks at ~ 783.0 , ~ 780.5 and ~ 778.2 eV indicate the presence of Co in the

oxidation states of Co^{2+} , Co^{3+} and zero-valent Co^0 , respectively.^{65,66} The two characteristic peaks at ~ 786.3 and ~ 803.8 eV represent the satellite peaks.⁶⁷ No traces of Zn were observed in the XPS analysis, which is expected, as the high-temperature pyrolysis has most likely evaporated the Zn from the carbon network.

2.2. Electrochemical performance of ZIF-8@CNT catalysts

A clear difference in the surface morphologies of the as-prepared ZIF-8@CNT catalysts can be observed from the results discussed in the earlier sections. These structural/morphological differences will affect the mass/charge transfer and total electrocatalytically active sites, which influence their ORR performance.⁶⁸ The electrochemical performance of the ZIF-8@CNT catalysts was investigated using a rotating disk electrode (RDE) setup. The linear sweep voltammetry (LSV) and cyclic voltammetry (CV) curves obtained for all the catalysts are displayed in Fig. 6. The ZIF-8@CNT catalysts showed excellent ORR performance at a low catalyst loading (0.1 mg cm^{-2}) in 0.1 M KOH solution (Fig. 6a). The metal-free catalyst (ZNT-900) showed the poorest ORR activity, which was expected due to its lowest BET surface area and absence of M-N_x centres. A clear influence of transition metal doping on the ORR performance can be observed in both single- and





Fig. 5 (a–e) Comparison of N 1s XPS spectra for (a) ZNT-900, (b) Fe-ZNT-900, (c) Co-ZNT-900, (d) Fe₁Co₁-ZNT-900 and (e) Fe₁Co₂-ZNT-900 catalysts and (f) N content for different nitrogen groups.

binary-metal-doped ZIF-8@CNT catalysts. Bimetallic ZIF-8@CNT catalysts (Fe₁Co₁-ZNT-900 and Fe₁Co₂-ZNT-900) showed the highest onset (at a current density $j = -0.1$ mA cm⁻²) and half-wave potentials (at $j = -3.0$ mA cm⁻²) of ~ 0.995 , ~ 0.996 V and 0.846 , 0.847 V vs. RHE, respectively, which surpass that of the commercially used Pt/C catalyst in terms of half-wave potential ($E_{1/2} = 0.834$ V). Both bimetallic ZIF-8@CNT catalysts reached similar limiting current densities (j_L) of -5.92 and -5.79 mA cm⁻², which is slightly lower compared to that of the Pt/C catalyst (-6.4 mA cm⁻²). This could be due to the well-developed micro-mesoporous

structure and a high number of new active sites derived from nitrogen and Fe/Co doping (M-N_x centres). Also, the presence of FeCo particles introduces high-density active centres, which makes them more accessible to the reactants to promote the ORR kinetics.⁶⁹ The ORR parameters obtained for all ZIF-8@CNT catalysts are listed in Table 3.

To further evaluate the electrocatalytic ORR behaviour of the catalysts, hydrodynamic voltammograms were recorded at different electrode rotation rates ($\omega = 360$ – 4600 rpm) as shown in Fig. S7.† To evaluate the oxygen reduction pathway, the electron transfer number (n) was calculated from



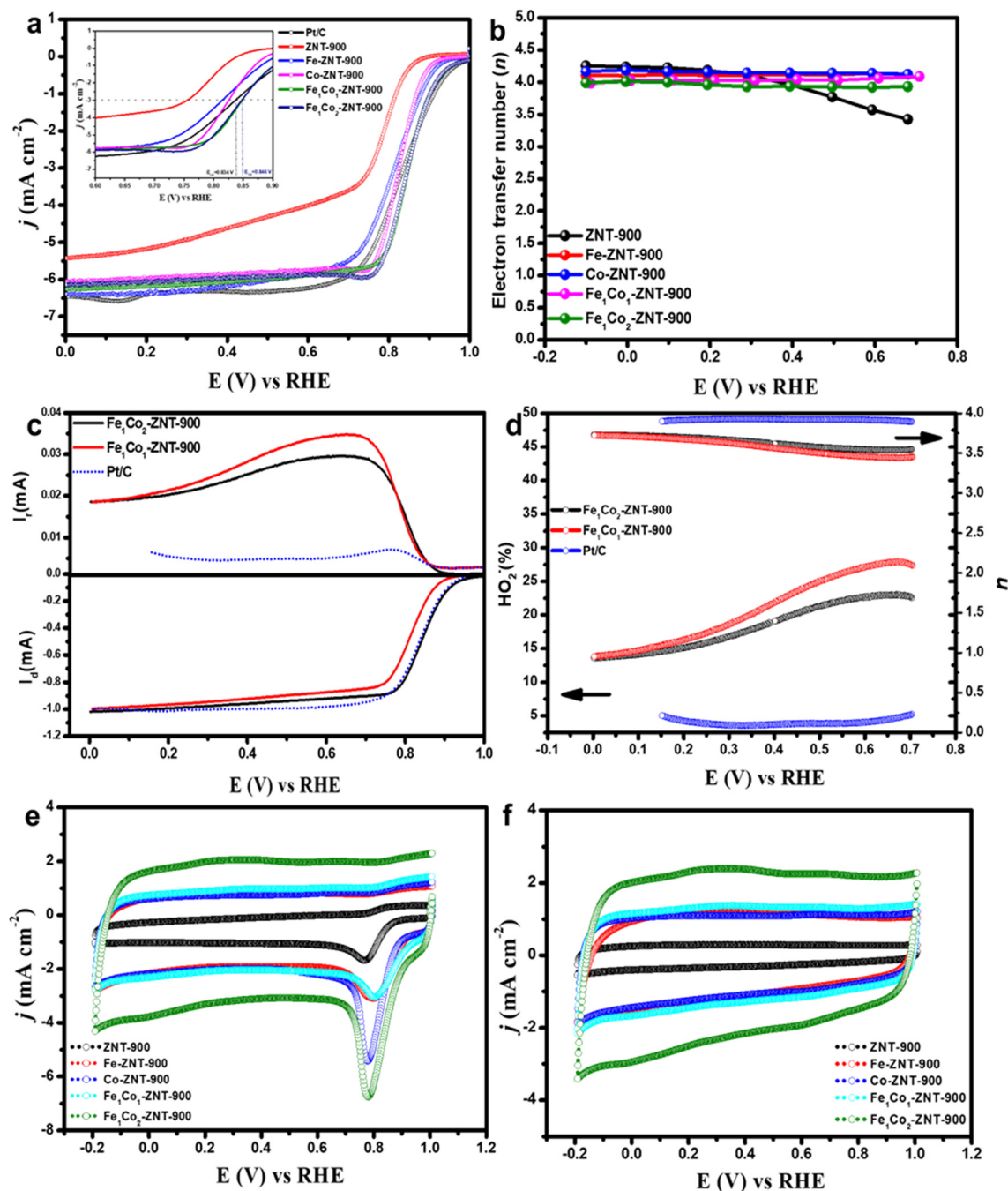


Fig. 6 (a) ORR polarisation curves for ZIF@CNT catalysts recorded at 1900 rpm with a scan rate of 10 mV s^{-1} in O₂-saturated 0.1 M KOH. (b) Electron transfer number (n) obtained from K-L plots for all the catalysts. (c) RRDE results for Fe₁Co₂-ZNT-900, Fe₁Co₁-ZNT-900 and Pt/C catalysts in 0.1 M KOH at 1900 rpm with a scan rate of 10 mV s^{-1} . (d) Electron transfer number (n) and HO₂[−] yield for Fe₁Co₂-ZNT-900, Fe₁Co₁-ZNT-900 and Pt/C catalysts. CV curves were recorded in (e) O₂-saturated and (f) Ar-saturated 0.1 M KOH at 100 mV s^{-1} .

Table 3 Summary of the half-wave potential ($E_{1/2}$), onset potential (E_{onset}), limiting current density (j_L), peak potential (E_p), and electron transfer number (n) values obtained for the ZIF-8@CNT catalysts

Catalyst	$E_{1/2}$ (V)	E_{onset} (V)	j_L (mA cm ^{−2})	E_p (V)	n
ZNT-900	0.750	0.874	−5.54	0.73	3.97 ± 0.31
Fe-ZNT-900	0.813	0.944	−6.27	0.79	4.0 ± 0.06
Co-ZNT-900	0.825	0.924	−5.93	0.80	4.05 ± 0.02
Fe ₁ Co ₁ -ZNT-900	0.846	0.995	−5.92	0.77	4.03 ± 0.03
Fe ₁ Co ₂ -ZNT-900	0.847	0.996	−5.79	0.78	3.96 ± 0.03

Koutecky-Levich (K-L) plots. In Fig. 6b, the number of electrons transferred during the ORR at the surface of ZIF-8@CNT catalysts is displayed and a clear influence of transition metal doping compared to the metal-free ZIF-8@CNT catalysts can be seen. The n value for ZNT-900 is 3.97 ± 0.31 and when Fe is doped the value of n reached 4.0 ± 0.06 . Also, for the bimetallic ZIF-8@CNT catalysts the n value was 4.03 ± 0.03 (Fe₁Co₁-ZNT-900) and 3.96 ± 0.03 (Fe₁Co₂-ZNT-900). To further study the ORR pathway, rotating ring-



disk electrode (RRDE) measurements were conducted to analyse the formation of hydrogen peroxide, which in turn were used to calculate the number of electrons transferred (n) (Fig. 6c). The collected ring (I_r) and disk currents (I_d) recorded for Fe₁Co₂-ZNT-900, Fe₁Co₁-ZNT-900 and Pt/C catalysts are presented in Fig. 6c. The RRDE polarisation curve presented in Fig. 6c (bottom) clearly shows that the bimetallic catalyst (Fe₁Co₂-ZNT-900) outperforms the Fe₁Co₁-ZNT-900 and Pt/C catalysts, which follows a similar trend observed in RDE polarisation curves presented in Fig. 6a. The HO₂⁻ yield for Fe₁Co₂-ZNT-900 reaches a maximum of 22.5% (at $E = 0.70$ V) and a minimum of 13.5%, delivering an average electron transfer number (n_{avg}) of 3.73, which is more than that of the Fe₁Co₁-ZNT-900 catalyst ($n_{\text{avg}} = 3.6$) as shown in Fig. 6d. The highly micro-mesoporous morphology of mono- and bimetallic ZIF-8@CNT catalysts prevented the aggregation of metals during the electrocatalytic process *via* stabilising the most accessible M-N_x active sites, which in turn accelerated the electron transfer process. Therefore, due to the presence of M-N_x active centres supported by XPS analysis, the transition-metal-doped electrocatalysts delivered the highest ORR performance when compared to the metal-free ZIF-8@CNT catalyst.

Fig. 6e and f represent the CVs recorded in O₂- and Ar-saturated 0.1 M KOH at the scan rate of 100 mV s⁻¹, respectively. The ORR peak potentials (E_p) for all the catalysts are listed in Table 3. Metal-free ZIF-8@CNT catalysts showed the lowest E_p , which can be attributed to their textural properties and absence of M-N_x centres. It can be observed that introducing transition metals into the ZIF-8@CNT 3D network creates more accessible M-N_x centres that accelerate the electron transfer process.⁷⁰ The Fe₁Co₂-ZNT-900 catalyst showed an intensive ORR peak at around 0.78 V vs. RHE due to the collective contribution of FeCo alloys and their textural properties. A similar behaviour can be seen in CV curves obtained in Ar-saturated 0.1 M KOH, with Fe₁Co₂-ZNT-900 and Fe₁Co₁-ZNT-900 catalysts carrying the highest capacitive currents in comparison to other catalysts. It is well evident that capacitive currents are highly dependent on the porosity of the catalyst material.⁷¹ Therefore, the higher capacitive currents for Fe₁Co₂-ZNT-900 and Fe₁Co₁-ZNT-900 are most likely due to their highly dense micro-mesoporous morphology with promising textural features. Based on the electrocatalytic activity of the prepared catalysts, Fe₁Co₂-ZNT-900 and Fe₁Co₁-ZNT-900 catalysts showed the maximum ORR application potential in terms of half-wave potential, onset potential, and their highly porous 3D ZIF structure with the presence of active N-centers and high number of M-N_x sites promoting the 4e⁻ ORR pathway. Therefore, these two catalysts were selected for further evaluation under the AEMFC conditions.

2.3. AEMFC performance

The AEMFC testing was performed with Pt-Ru/C and FeCo-ZNT-900 (Fe:Co = 1:1 and 1:2) catalysts used at the anode

and cathode, respectively. An open circuit voltage (OCV) of 1.0, 1.03, and 1.01 V was recorded for Fe₁Co₁-ZNT-900-, Fe₁Co₂-ZNT-900-, and Pt/C-coated cathodes, respectively, as shown in the j - V polarisation curves (Fig. 7a). This indicates the better intrinsic electrocatalytic activity of these FeCo-based catalysts which in turn is accredited to the highly accessible active centres.⁷² The power density curves shown in Fig. 7b indicate that the Fe₁Co₂-ZNT-900 catalyst performs better than Fe₁Co₁-ZNT-900 with a maximum power density (P_{max}) of 0.171 W cm⁻², which is comparable to that of the commercial Pt/C catalyst (0.215 W cm⁻²). Moreover, the current density of the Fe₁Co₂-ZNT-900 catalyst at 0.5 V ($j_{0.5}$) was 0.326 A cm⁻², which is comparable to the that of the Pt/C catalyst ($j_{0.5} = 0.359$ A cm⁻²). Also, it is evident from the RRDE results that the Fe₁Co₂-ZNT-900 catalyst showed about 17.6% lower HO₂⁻ yield when compared to the Fe₁Co₁-ZNT-900 catalyst, which affects the overall ORR kinetics and pathway that is reflected by the better AEMFC performance of the Fe₁Co₂-ZNT-900 catalyst. Most importantly, in the voltage range considered suitable for practical AEMFC use (between 0.8 and 0.6 V),^{73,74} Fe₁Co₂-ZNT-900 outperforms the Pt/C cathode. This is ascribed to the effective mass transport of reactants at the active sites in the electrodes due to the hierarchical porous texture in the catalyst materials. It should be noted that the ORR performance of the catalysts evaluated from RDE does not provide a realistic picture when their activity in the fuel cell is taken into consideration.⁷⁵ The probable reasons for such discrepancies can be ascribed to varying operational conditions in fuel cells such as mass transport, operational temperature, water transfer regulations, most importantly the insufficient ionic conductivity of AEM, *etc.* Recent reports suggest that the adjustment of the catalyst layer with the ionomer and AEM plays a vital role in generating promising power densities.^{74,76} Therefore, the ORR activity of the catalysts in 0.1 M KOH solution and their corresponding AEMFC performance are not comparable. In previous studies, we have tested Fe- and Co-based catalysts in an AEMFC using a 10 μm thick Aemion + AEM, but no reports are available with the Aemion+@ 15 μm AEM.^{38,75,77–80} The Aemion+ AEM (10 μm) consists of repeating units of methylated polybenzimidazoles, which offers high stability in AEMFCs by employing the steric-crowding strategy to stabilise the C-2 position of the imidazole group.⁸¹ Also, the thickness of the AEM plays a crucial role in the efficient migration of OH⁻ ions and H₂O molecules.⁸² Wei *et al.* demonstrated that the higher thickness of the Aemion+ AEM (25 μm) exhibits lower AEMFC performance by achieving 23.5% lower P_{max} in comparison to 10 μm Aemion AEM.⁸³ A thinner AEM provides better water back-diffusion from the membrane, resulting in more favourable water management. In addition to membrane thickness, the ion exchange capacity (IEC) also affects the AEMFC performance. An AEM with higher thickness possesses lower IEC, which as a result increases OH⁻ ionic resistance and lowers fuel cell performance.⁸⁴ Aemion+ (10 μm) exhibits a higher IEC of 3.0 mmol g⁻¹, which is 33.3%





Fig. 7 (a) Polarisation and (b) power density curves obtained for Fe₁Co₁-ZNT-900, Fe₁Co₂-ZNT-900, and Pt/C cathode catalysts in H₂/O₂ AEMFCs [anode catalyst: PtRu/C. AEM: Aemion+® 15 μ m. $T = 65$ °C].

more than for the Aemion+ (15 μ m) AEM, resulting in higher ionic resistance in the case of Aemion+® 15 μ m.⁸⁵ The latter can be observed by considerably lower power output in the case of Aemion+® 15 μ m in comparison with previously reported Aemion+ AEMs having a thickness of 10 μ m. Despite the similar building blocks in Aemion+ and Aemion+® 15 μ m, the noticeable variation in their physical and chemical properties resulted in the performance difference in AEMFC. The peak power densities obtained for FeCo-based catalysts in this work are not very high as compared to those of similar catalyst materials reported in the literature. In fact, considering all the influencing factors in AEMFC operation and especially the AEM, the catalysts performed quite well in fuel cell conditions and further optimisation is being conducted to achieve higher power densities.

3. Conclusions

Metal-free and transition-metal-doped ZIF-8@CNT catalysts were prepared by one-step high-temperature pyrolysis at 900 °C. The existence of M-N_x sites was shown by XPS analysis and XRD analysis showed the presence of pure Fe and Co nanoparticles and FeCo nanoalloys with a minimal existence of Fe₃O₄ nanoparticles. The presence of alloys was further confirmed by using STEM measurements. All the transition-metal-doped ZIF-8@CNT catalysts showed good ORR activity, with metal-free ZNT-900 being the poorer-performing electrocatalyst among the tested ones. The ZIF-8@CNT catalysts showed excellent electrocatalytic ORR activity in 0.1 M KOH at a low catalyst loading of 0.1 mg cm⁻². Bimetallic ZIF-8@CNT electrocatalysts (Fe₁Co₂-ZNT-900) delivered the highest half-wave and onset potential of 0.847 and 0.996 V vs. RHE, respectively, and outperformed the Pt/C catalyst in terms of half-wave potential (0.834 V). The performance of bimetallic ZIF-8@CNT electrocatalysts was investigated in AEMFCs, and the Fe₁Co₂-ZNT-900 cathode reached a P_{\max} of 0.171 W cm⁻², which was close to the P_{\max} achieved with the Pt/C catalyst. The AEMFC results can be improved by further optimising various

factors, such as the AEM, operational temperature, water management, etc. All in all, the bimetallic ZIF-8@CNT catalysts performed quite well in AEMFCs and the results were comparable to those of similar materials and these catalysts should be explored more in-depth to achieve more promising results in AEM fuel cells.

4. Experimental section

4.1. Chemicals

Commercially available zeolitic imidazolate framework-8 (ZIF-8) (Basolite® Z1200, Sigma-Aldrich), multiwalled carbon nanotubes (MWCNTs, NC3150, $\geq 95\%$, Nanocyl S.A.), cobalt(II) acetate (Co(OAc)₂, 98%, Alfa Aesar), iron(II) acetate (Fe(OAc)₂, 95%, Sigma-Aldrich), Nafion solution (5 wt% in lower aliphatic alcohols, Sigma-Aldrich), 2-propanol (Honeywell), methanol (MeOH, 99.99%, Lach-Ner), and potassium hydroxide (KOH, $\geq 85\%$, Sigma-Aldrich) were used without further purification. Milli-Q water was used to prepare the solutions.

4.2. Synthesis of ZIF-8@CNT catalysts

The catalyst materials were prepared by adding different proportions of Fe and Co salts with a fixed amount of ZIF-8 and MWCNTs as summarised in Table S1.† First, a catalyst without metal addition was prepared by mixing 105 mg of ZIF-8 and 100 mg of MWCNTs in 40 mL MeOH solvent and named ZNT-900 after the pyrolysis of the dried mixture. Fe-doped catalyst was prepared by mixing 8.2 mg of Fe(OAc)₂, 100 mg of MWCNTs, and 105 mg of ZIF-8 in MeOH. The amount of Fe was 2.5 wt% of MWCNTs and was selected from the procedure mentioned by Ratso *et al.*⁸⁶ Similarly, all other catalyst dispersions were prepared and sonicated (Branson 1510E-MTH; power output = 70 W, frequency = 40 kHz, Branson) for about 1.5–2 h to obtain a homogenous slurry followed by oven-drying at 60 °C overnight. The blackish powder obtained after drying was then pyrolysed at



900 °C for 2 h at a ramp rate of 10 °C min⁻¹ in an inert atmosphere maintained by the continuous supply of N₂ gas throughout the pyrolysis tube (N₂, 99.99%, Linde Gas). In all, five different catalysts were prepared and were named as ZNT-900, Fe-ZNT-900, Co-ZNT-900, Fe₁Co₂-ZNT-900 (Fe:Co = 1:2), and Fe₁Co₁-ZNT-900 (Fe:Co = 1:1). If not stated otherwise, Z stands for ZIF-8, NT for MWCNTs and 900 represents the pyrolysis temperature of 900 °C.

4.3. Physicochemical characterisation of ZIF-8@CNT catalysts

The morphology of the catalyst materials was examined by high-resolution scanning electron microscopy (HR-SEM, Helios NanoLab 600, FEI Company) at 20 kV and scanning transmission electron microscopy (STEM, Titan Themis 200 (FEI) (S)TEM) at 200 kV. An energy-dispersive X-ray (EDX) spectrometer combined with SEM and STEM instruments (INCA Energy 350, Oxford Instruments, Abingdon, UK, and Super-X EDX, FEI/Bruker, respectively) was used to analyse the elemental composition of the catalyst materials. An X-ray diffractometer (XRD, Bruker D8 Advance), supplied with Cu K α as a radiation source and a LynxEye line detector, and an X-ray photoelectron spectrometer (XPS, Scienta SES-100 electron analyser using Mg K α radiation from a Thermo XR3E2 X-ray source) were used to study the crystalline structure and surface chemical composition of the catalyst materials, respectively. The specific surface area (S_{BET}) and porosity measurements were calculated using the Brunauer–Emmett–Teller (BET) model and N₂ adsorption–desorption isotherms, respectively. The total pore volume (V_{tot}) was measured at a N₂ saturation pressure (P/P_0) of 0.97. The pore size distribution (PSD) was estimated from N₂ isotherms using a quenched solid density functional theory (QSDFT) equilibrium model for slit-type pores. Micro-Raman spectra were recorded in back-scattering geometry on an inVia Renishaw spectrometer in conjunction with a confocal microscope (Leica Microsystems CMS GmbH, 50 \times objective) and an argon ion laser operated at 514.5 nm. Low incident laser power density at the sample prevented excessive sample heating and/or decomposition. The metal contents in the catalyst materials were analysed by microwave plasma atomic emission spectroscopy (MP-AES) using an Agilent 4210 MP-AES instrument.

4.4. Electrochemical studies of ZIF-8@CNT catalysts

All the electrochemical measurements were performed on catalyst-coated glassy carbon (GC) disks (GC-20SS, Tokai Carbon Ltd., Japan). The GC disks inserted into a Teflon holder were used as electrode substrates for ORR studies. To clean the GC surface, it was polished using P1200 SiC abrasive paper and 1 and 0.3 μm alumina slurries (Buehler, USA) until a shiny and reflective mirror-like appearance was observed. The GC electrodes were further cleaned *via* sonication in Milli-Q water and 2-propanol for 5 min in each solvent. The catalyst ink was prepared by suspending 8 mg of the catalyst powder in 2 mL of 2-propanol and 10 μL of 5

wt% Nafion ionomer solution followed by sonication until a homogenous suspension was obtained. A certain aliquot (1.7 μL) of the catalyst suspension was pipetted onto the GC surface and allowed to dry in air. A total of 5 μL catalyst ink was deposited on the pre-polished GC surface to obtain a catalyst loading of 0.1 mg cm⁻².

A standard round-shaped five-neck glass cell (volume = 100 mL) with a three-electrode configuration was used to measure the electrochemical performance of the as-prepared ZIF-8@CNT catalysts in 0.1 M KOH electrolyte solution (pH \sim 13) using an Autolab potentiostat/galvanostat PGSTAT128N (Metrohm Autolab, the Netherlands). The catalyst-coated GC electrode, saturated calomel electrode (SCE), and a carbon rod were used as the working, reference, and counter electrode, respectively. The electrode rotation rate (ω = 360–4600 rpm) was controlled using an EDI101 rotator and CTV101 speed control unit (Radiometer). Rotating disk electrode (RDE) voltammetry and cyclic voltammetry (CV) curves were recorded in the potential window of -1.2 to 0.0 V (*vs.* SCE) at a scan rate (v) of 10 mV s⁻¹; later, the recorded potentials were converted into potential values *vs.* reversible hydrogen electrode (RHE) using the equation $E_{\text{RHE}} = E_{\text{SCE}} + 0.241 + 0.059 \times \text{pH}$.⁸⁷ A commercial Pt/C (20 wt% Pt) catalyst was used for comparing the electrochemical performance of the as-prepared ZIF-8@CNT catalysts. All the RDE voltammetry curves were IR-compensated using ohmic resistance values obtained with electrochemical impedance spectroscopy. If not stated otherwise, all the electrochemical experiments were conducted in O₂- or Ar-saturated (99.999%, Linde) 0.1 M KOH electrolyte. To evaluate the ORR pathway, electron transfer number (n) was calculated from Koutecky–Levich (K–L) plots using eqn (1)–(3).^{39,88}

$$j^{-1} = j_{\text{k}}^{-1} + B^{-1}\omega^{1/2} \quad (1)$$

$$j_{\text{k}} = nFkC_{\text{O}} \quad (2)$$

$$B = 0.62 nFC_{\text{O}}(D_{\text{O}})^{2/3}v^{-1/6} \quad (3)$$

where j and j_{k} represent the experimental and kinetic current densities at a specific potential, j_{d} is the diffusion-limited current density, n is the number of electrons transferred per O₂ molecule, k is the O₂ reduction rate constant at a specific potential (cm s⁻¹), F is the Faraday constant (96485 C mol⁻¹), D_{O} is the O₂ diffusion coefficient in 0.1 M KOH (1.9×10^{-5} cm² s⁻¹), v is the kinematic viscosity of the electrolyte solution (0.01 cm² s⁻¹), C_{O} is the concentration of O₂ in the bulk solution (1.2×10^{-6} mol cm⁻³) and ω is the electrode rotation rate (rad s⁻¹).⁸⁹

To further investigate the ORR pathway, the rotating ring-disk electrode (RRDE) measurements were conducted using an MSR speed controller and an AFMSRX rotator (Pine Research, USA) to evaluate the hydroperoxide anion (HO₂⁻) yield (%) and the number of electrons transferred (n).⁹⁰ A fixed-disk RRDE tip composed of a GC disc-Pt ring was attached to the AFMSRX rotator. The GC projected a



geometric surface area of 0.164 cm^2 with a catalyst loading of 0.1 mg cm^{-2} . The catalyst inks were prepared similarly to RDE measurements. The experimental data were recorded using General Purpose Electrochemical System (GPES) software. The $\% \text{HO}_2^-$ and n values were calculated using eqn (4) and (5), respectively.

$$\% \text{HO}_2^- = \frac{2I_r/N}{I_d + I_r/N} \times 100\% \quad (4)$$

$$n = \frac{4I_d}{I_d + I_r/N} \quad (5)$$

The detection of HO_2^- was measured using a Pt ring electrode with a collection efficiency (N) of 0.25 at an applied potential of 1.55 V vs. RHE. The Pt ring was activated before use by running at least three potential cycles within a 0.05–1.55 V (vs. RHE) window at the electrode rotation rate of 1900 rpm with a scan rate of 100 mV s^{-1} .

4.5. MEA preparation and AEMFC tests

The membrane-electrode assembly (MEA) was prepared by sandwiching ZIF-8@CNT catalyst-coated gas diffusion layer (GDL, Sigracet 39 BB) with Aemion+® 15 μm (AF2-HLF8-15-X, 15 μm thickness, Ionomr Innovations, Canada) anion-exchange membrane (AEM) as shown in Fig. S1.† Cathode inks were prepared by dispersing 12.5 mg of FeCo-ZNT-900 in 1350 μL of methanol, 390 μL of Milli-Q water and 130 μL of 3 wt% ionomer solution prepared *via* dissolving Aemion+ powder (AP2-INN8-00-X) (Ionomr Innovations, Inc.) in methanol. The catalyst ink for the anode was prepared by dispersing 6.67 mg of Pt-Ru/C (50:25:25, Alfa Aesar), 690 μL of methanol, 196 μL of Milli-Q water, and 49 μL of 3 wt% same ionomer solution. The homogeneously dispersed ink was obtained by sonication of the suspensions under an ice bath for about 90 min. Both the prepared anode and cathode inks were incrementally deposited onto the gas diffusion layers until a loading of $0.8 \text{ mg}_{\text{Pt-Ru}} \text{ cm}^{-2}$ and 2 mg cm^{-2} , respectively, were achieved. To ensure sufficient ion exchange, the as-prepared electrodes and anion-exchange membrane were soaked in 3 M KOH solution for 24 and 96 h, respectively, prior to assembling the AEMFC setup. The 3 M KOH solution was frequently replenished every 24 h. The pretreated AEM, electrodes, and silicone gaskets were fixed in a single-cell fixture (Fuel Cell Technologies Inc., USA) of 5 cm^2 active area at 9 Nm torque. Single-cell AEMFC experiments were conducted using a Greenlight fuel cell test station (G40 Fuel Cell System, Hydrogenics, Canada) by feeding 65% humidified H_2 and O_2 gases at a flow rate of 1 NLPM at a backpressure of 0.2 MPa; the cell temperature was held at 65°C .

Conflicts of interest

The authors declare no conflict of interest.

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