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Engineering crystal structure of Cs₃BiBr₆ and Cs₃Bi₂Br₉ is theoretically and experimentally demonstrated to modulate their photocatalytic performance. This work offers insights into the structure-photoactivity relationships of metal halide perovskites (MHPs) and provides a guideline for exploiting MHPs toward efficient photocatalytic organic synthesis.

Semiconductor-based photocatalysis has attracted growing attention in solar-to-chemical energy conversion.1 The basic principle of photocatalysis starts at the electronic band structure of the semiconductors, which dominates the optoelectronic properties and drives the charge carrier that underpin their photocatalytic properties.² Consequently, the semiconductor crystal structure, which determines the electronic band structures, as well as the photogenerated charge carrier dynamics, is a key factor affecting their photocatalytic performance.³ For example, anatase TiO₂ has often been reported to exhibit better photoactivity than the rutile crystal phase, originating from the difference in photogenerated charge generation and transfer.^{4,5} Similarly, tetragonal BiVO₄ was reported to be less active as a photocatalyst than monoclinic scheelite BiVO₄.6 In short, the photocatalytic performance of semiconductors can be regulated via crystal structure engineering.

Inspired by the great advances in solar cells, ^{7,8} metal halide perovskites (MHPs) have recently emerged as promising photocatalysts due to their excellent optoelectronic properties. 9-11

Among them, lead-free Bi-based compounds such as Cs₃Bi₂Br₉ show excellent stability.12 However, most MHP materials generally exhibit modest photocatalytic performance due to ineffective charge separation and sluggish surface reaction. To improve photoactivity, cocatalyst loading^{9,13} and heterojunction construction 10,11 are effective. On the other hand, the crystalline structure of MHPs has been shown to play a key role in modulating photoactivity. 13 For instance, the incorporation of heteroatoms Ag13 and Sb14 into Cs3Bi2Br9 leads to form Cs₂AgBiBr₆ and Cs₃SbBiBr₉, which could improve the photoactivity as compared with Cs3Bi2Br9. However, Cs2AgBiBr6 and Cs₃SbBiBr₉ are intrinsic different from Cs₃Bi₂Br₉, making it difficult to directly uncover the relationship between crystal structures and photocatalytic performance of MHPs. Therefore, it is anticipated to explore the exact effect of crystal structure on photoactivity based on the MHPs with same elements while different crystalline structures.

Herein, the effect of crystal structure in Cs₃BiBr₆ and Cs₃Bi₂Br₉ on the photocatalytic performance is investigated. Density functional theory (DFT) calculations reveal that the photogenerated charge carriers in Cs₃Bi₂Br₉ have a reduced effective mass, lower exciton binding energy, and longer carrier lifetime than that in Cs₃BiBr₆. Photoelectrochemical characterizations further confirm that Cs3Bi2Br9 exhibits improved charge transfer efficiency as compared with Cs₃BiBr₆. When applied to photocatalytic benzyl alcohol (BA) oxidation, Cs₃Bi₂Br₉ displays enhanced photocatalytic performance compared to Cs₃BiBr₆. Further decoration of these MHPs with Pd nanocubes as cocatalysts, the BA conversion over 1 wt% Pd/ Cs₃Bi₂Br₉ (82.3%) is over twice than that of 1 wt% Pd/Cs₃BiBr₆ (39.5%). This work offers mechanistic insights into the structure-photoactivity of MHPs and provides a guideline for rational design of efficient MHP photocatalysts.

DFT calculations were first performed to understand the crystal structures of Cs₃Bi₂Br₉ and Cs₃BiBr₆. For Cs₃Bi₂Br₉, within the unit cell (Fig. 1A, Table S1, ESI†), eight Cs⁺ ions occupied the corners and two other Cs⁺ ions located on the body diagonal of the unit cell, 13 and Bi atoms occupy partial Br₆

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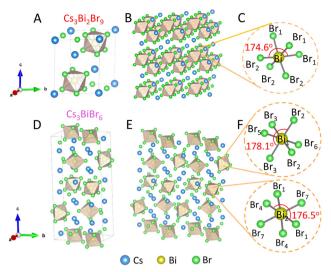


Fig. 1 (A) The unit cell of Cs₃Bi₂Br₉. (B) Three-dimensional structure of Cs₃Bi₂Br₉ material. (C) Bi-Br bonds and the angles of Br₁-Bi-Br₂ of BiBr₆ octahedra in Cs₃Bi₂Br₉. (D) The unit cell of Cs₃BiBr₆ material. (E) Threedimensional structure of Cs₃BiBr₆. (F) Bi-Br bonds and the typical Br₂-Bi₁-Br₃ and Br₇-Bi₂-Br₇ angles of BiBr₆ octahedra in Cs₃BiBr₆.

holes in an ordered arrangement (Fig. 1B), where three short and three long Bi-Br bonds with shorter linkages are shared between neighboring octahedra in each layer (Fig. 1C, Table S2, ESI†).¹³ The local coordination environment of the [BiBr₆]³⁻ octahedra is irregular and leads to a trigonal distortion, forming a two-dimensional perovskite structure. Cs₃BiBr₆, by contrast, contains two Bi atoms, three Cs atoms, and seven Br atoms in an asymmetric unit (Fig. 1D). These two Bi atoms are independent (Fig. 1F), where each Bi atom is coordinated by six Br atoms and the bond lengths of Bi-Br are different (Table S2, ESI†), and the bond angle between Br-Bi-Br in Cs₃BiBr₆ is distorted.15 The [BiBr₆]3- octahedra exist in isolation and no link between these octahedra, forming a zero-dimensional perovskite structure (Fig. 1E). 15,16 Then, the band structures and density of states (DOS) of both materials were studied. From the energy band structures (Fig. S1, ESI†), the conduction band minimum (CBM) and valence band maximum (VBM) of Cs₃Bi₂Br₉ are more dispersive than those of Cs₃BiBr₆, hinting that the Cs₃Bi₂Br₉ has lower effective masses of electrons and holes (vide infra). 13 For Cs₃Bi₂Br₉, the VBM primarily consists of Br 4p states and Bi 6p orbitals, mixed with a few Bi 6s states (Fig. S1B, ESI†), and the CBM is dominated by Bi 6p and Br 4p orbital interactions.¹⁷ For Cs₃BiBr₆, the VBM is mainly composed of Br 4p orbitals, and the CBM consists of an admixture of Bi 6p and Br 4p orbitals (Fig. S1D, ESI†).

The photoelectric properties of Cs₃Bi₂Br₉ and Cs₃BiBr₆ can be assessed by physical factors such as exciton binding energy, effective mass, charge mobility, and charge carrier relaxation time. 13 These parameters were calculated and listed in Table 1. The exciton binding energy (E_b) of $Cs_3Bi_2Br_9$ is calculated to be 0.278 eV, much smaller than that of Cs₃BiBr₆ (0.505 eV), which is beneficial for the separation of photogenerated excitons and extraction of charge carriers. 13 Besides, Cs₃Bi₂Br₉ shows a

Table 1 Calculated electron and hole effective masses (m^*) , carrier mobility (μ) , corresponding relaxation time (τ) , exciton binding energy (E_b) , and bandgap (E_a) for $Cs_3Bi_2Br_9$ and Cs_3BiBr_6

Parameter	$\mathrm{Cs_3Bi_2Br_9}$	$\mathrm{Cs_3BiBr_6}$
$m_{\rm e}^* (m_{\rm e})$	0.57	1.42
$m_{\rm h}^* (m_{\rm h})$	5.46	6.32
$\mu_{\rm e} (m^2 {\rm V}^{-1} {\rm S}^{-1})$	4.89×10^{-1}	6.34×10^{-2}
$\tau_{\rm e}$ (fs)	1585.3	512.03
$E_{\rm b}$ (eV)	0.278	0.505
$E_{\rm g}$ (eV)	2.596	3.195

significant drop in effective mass of charge carriers as compared to Cs_3BiBr_6 ; calculated electron effective mass 0.57 m_e $(Cs_3Bi_2Br_9)$ vs. 1.42 m_e (Cs_3BiBr_6) and the effective hole mass reduces from 6.32 $m_{\rm h}$ to 5.46 $m_{\rm h}$. Generally, the smaller the effective mass of charge carriers, the faster the transfer rate of charges. 10 Thus, the lower effective mass in Cs₃Bi₂Br₉ links to higher mobility which is beneficial for charge carriers migration.¹³ This is confirmed by the calculated carrier mobility, where the value of Cs3Bi2Br9 is higher than that of Cs₃BiBr₆. Also, their different deformation potential constants (Fig. S2, ESI†) agree with the results of effective mass and carrier mobility. Furthermore, Cs3Bi2Br9 possesses a longer carrier relaxation time than Cs3BiBr6, which responds to the prolonged carrier lifetime in Cs₃Bi₂Br₉.

Next, Cs₃Bi₂Br₉ and Cs₃BiBr₆ were synthesized (see ESI†) for further analysis. 18 XRD results in Fig. 2A show that both Cs₃Bi₂Br₉ and Cs₃BiBr₆ exhibit strong diffraction peaks, without shift compared to theoretical data. Cs3Bi2Br9 has a trigonal crystalline structure, 19 while Cs₃BiBr₆ crystallizes in an orthorhombic phase. 15 Raman spectroscopy (Fig. 2B) reveals the modes at 167 and 192 cm⁻¹ in Cs₃Bi₂Br₉ belonging to the stretching vibrations of Bi-Br bonds in the [BiBr₆]³⁻ octahedra.²⁰ For Cs₃BiBr₆, although

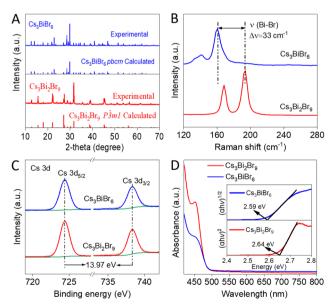


Fig. 2 (A) XRD patterns, (B) Raman spectra, (C) high-resolution XPS spectra of Cs 3d core levels, and (D) DRS spectra of Cs₃Bi₂Br₉ and Cs₃BiBr₆ materials. The inset in Fig. 2D shows the bandgaps of both materials.

it also contains $[BiBr_6]^{3-}$ octahedra, the complex groups are isolated and do not share corners, yielding peaks at different frequencies, *i.e.* 160 and 135 cm⁻¹. By contrast, Cs_3BiBr_6 showed a red-shift (33 cm⁻¹) for the terminal Bi–Br stretching modes of the $[BiBr_6]^{3-}$ octahedra compared to $Cs_3Bi_2Br_9$, which suggests weaker Bi–Br bonds in Cs_3BiBr_6 , ¹⁶ in line with the slightly longer Bi–Br bond lengths (Table S2, ESI†).

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XPS was performed to investigate the chemical structure and valence state of both materials. The signals of Cs 3d (Fig. 2C), Bi 4f and Br 3d (Fig. S3A and B, ESI†) are the same, confirming their expected valence states (+1, +3, and -1, respectively) to be consistent with the formula. ^{15,16} The stoichiometry in Cs₃Bi₂Br₉ was calculated to be Cs:Bi:Br = 3:2.21:9.20, and for Cs₃BiBr₆ the ratio is Cs:Bi:Br = 3:1.22:6.37 (Table S3, ESI†), close to their theoretical values. SEM images (Fig. S4, ESI†) show that both materials have a similar particle size of about 500 nm. EDX spectroscopy further confirms the presence of Cs, Bi and Br elements (Fig. S4E and F, ESI†). Nitrogen physisorption data (Fig. S5, ESI†) shows that the surface area of both materials is similar (Table S4, ESI†).

UV-Vis DRS was employed to determine their absorption properties (Fig. 2D). Note that although Cs₃Bi₂Br₉ and Cs₃BiBr₆ have different elemental compositions and structures, they have in common [BiBr₆]³⁻ octahedra, and their optical properties are largely determined by the Bi³⁺ cation levels, perturbed slightly by the environment surrounding [BiBr₆]³-octahedra.20,21 The absorption onset of Cs3Bi2Br9 is measured at 470 nm (2.64 eV), ¹⁸ and for Cs₃BiBr₆ it is 479 nm (2.59 eV), ¹⁵ which is consistent with their colors (Fig. S6, ESI†). The extracted bandgap of Cs3Bi2Br9 agrees with DFT calculation as well as reported results, 18 while Cs₃BiBr₆ shows a difference (0.6 eV) with the calculated one which could be attributed to the DFT limitations. 15 Mott-Schottky analyses (Fig. S7, ESI†), combined with UV-Vis data, show that CB and VB edge potentials of $Cs_3Bi_2Br_9$ and Cs_3BiBr_6 are -0.57 and 2.07 eV, and -0.69 and 1.9 eV (vs. NHE), respectively. Steady-state PL spectrum (Fig. S8, ESI†) shows that Cs₃Bi₂Br₉ has an emission centered at around 469 nm, which almost coincides with its bandgap (2.64 eV), corresponding to near band edge emission. 13 The emission peak of Cs₃BiBr₆ shows a slight blue shift, which could be attributed to its indirect bandgap property.22

Photoelectrochemical measurements were further performed to reveal the picture of charge separation and transfer. ²³ Fig. 3A shows the transient photocurrent responses of both materials, where a higher photocurrent is observed for $Cs_3Bi_2Br_9$ than that of Cs_3BiBr_6 , suggesting the efficient charge separation in $Cs_3Bi_2Br_9$. ²⁴ To better understand the difference, EIS experiments were performed. ²³ Clearly, $Cs_3Bi_2Br_9$ presents a smaller semicircle in the Nyquist plots (Fig. S9, ESI†), confirming the efficient transfer of charge carriers in $Cs_3Bi_2Br_9$. ^{24–26} This is further corroborated by the surface charge transfer efficiency (η_t) measurements ^{25,27} (Fig. 3B), the η_t of $Cs_3Bi_2Br_9$ and $Cs_3Bi_2Br_9$ is calculated to be 24% and 31%, respectively. To conclude, the above observations disclose that $Cs_3Bi_2Br_9$ would empower faster charge carriers injection into the redox couple during the photocatalytic reaction. ^{10,24}

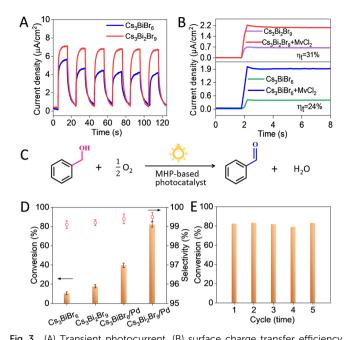


Fig. 3 (A) Transient photocurrent, (B) surface charge transfer efficiency (η_t) of Cs₃Bi₂Br₉ and Cs₃BiBr₆. (C) Schematic illustration of photocatalytic selective oxidation of benzyl alcohol (BA) to benzaldehyde (BAD) under simulated solar light illumination. (D) The photocatalytic performance of selective BA oxidation over various photocatalysts. (E) Stability test of Pd/Cs₃Bi₂Br₉ composite. Reaction conditions: 0.1 mmol BA in 2.5 mL molecular oxygen saturated benzotrifluoride, 15 mg catalyst, simulated solar light (AM 1.5G, 150 W Xe lamp), reaction time (6 h).

The selective photocatalytic oxidation of benzyl alcohol (BA) was used as model reaction to compare the performance of Cs₃Bi₂Br₉ and Cs₃BiBr₆ (Fig. 3C, see Scheme S1, ESI†); the band edge energies of both materials are suitable for the involved redox potentials (+ 0.68 eV). 28 Cs3Bi2Br9 shows a higher BA conversion (17.5%) than Cs₃BiBr₆ (10.2%) after solar light irradiation of 6 h (Fig. 3D), with an excellent selectivity (over 99%) towards benzaldehyde (BAD). To further enhance photoactivity, Pd nanocubes were deposited on MHPs surfaces. The 1 wt% Pd/Cs₃Bi₂Br₉ composite converts 82.3% BA, over 2-fold enhancement than 1 wt% Pd/Cs₃BiBr₆ (39.5%). This is attributed to the fact that Pd serves as an electron reservoir to promote the transfer of photogenerated charge carriers, thus improving the conversion of BA.27 Besides, a comparison of Cs₃Bi₂Br₉/Pd with reported photocatalysts for selective BA oxidation shows the excellent performance of this material (Table S5, ESI†). Moreover, recycling tests show that no activity loss for 1 wt% Pd/Cs₃Bi₂Br₉ is observed after 5 successive cycles (Fig. 3E). Additionally, the catalyst is stable after exposing in air for 15 days (Fig. S10, ESI†), and no crystal structure changes are observed in XRD patterns (Fig. S11, ESI†).

To elucidate the reation mechanism over Pd/Cs₃Bi₂Br₉ sample, control experiments using different radical scavengers were performed (Fig. S12, ESI†). When hole scavenger ammonium oxalate was added, the conversion rate shows a sharp drop, indicating the key role of photogenerated holes in driving this reaction.¹⁰ Contrarily, the performance remains with electron

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(potassium persulfate) or •O₂⁻ (1,4-benzoquinone) scavengers. ^{10,29} This follows that O₂ was trapped by electrons to produce •O₂⁻ radicals, while the generated ${}^{\bullet}O_2^{-}$ is not directly involved in the BA oxidation process.²³ Moreover, when the O₂ is replaced by argon (Ar), an obvious reduction in the activity was observed, showing that O₂ reacted with electrons which can effectively promote the separation of charge carriers so that this reaction can proceed smoothly. In addition, the presence of hydroxyl radicals (*OH⁻) scavenger, t-butanol, has a negligible effect on the performance since no OH existed in this reaction system. Based on these experiments, we propose a plausible mechanism (Fig. S13, Scheme S1, ESI†). Under light illumination, Cs₃Bi₂Br₉ absorbs solar light to produce electron-hole pairs. The photogenerated electrons migrate to the surface of Pd and reduce O2 to O2, and the generated holes activate BA and generate the BAD product.²³

In summary, the effect of the crystal structure of Cs₃BiBr₆ and Cs₃Bi₂Br₉ on optoelectronic properties and ultimately photocatalytic performance, was investigated theoretically and experimentally. DFT results showed that Cs₃Bi₂Br₉ has a smaller carrier effective mass, lower exciton binding energy, and higher charge mobility than Cs₃BiBr₆. Furthermore, the more efficiency in charge separation within the Cs₃Bi₂Br₉ than Cs₃BiBr₆ was experimentally confirmed. As a result, Cs₃Bi₂Br₉ material exhibited a boosted photocatalytic BA oxidation performance, over twice higher than Cs₃BiBr₆. This work systematically illustrates the crystal structure-photoactivity relationship of MHPs and offers a guideline for exploiting excellent MHP photocatalysts.

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Conflicts of interest

There are no conflicts to declare.

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