



Cite this: *Dalton Trans.*, 2016, **45**, 4060

Received 4th September 2015,
Accepted 13th October 2015

DOI: 10.1039/c5dt03443d

www.rsc.org/dalton

Neutral N-donor ligand based flexible metal–organic frameworks

Biplab Manna,[†] Aamod V. Desai[†] and Sujit K. Ghosh*

This short review gives a focussed discussion on metal–organic frameworks (MOFs) made of neutral N-donor ligands which show structural flexibility under various exogenous stimuli. Chemical stimuli such as presence of anions, free guests, coordinated guests and physical stimuli (light, heat and so on) render structural flexibility in MOFs. Single-crystal-to-single-crystal transformation studies have attracted a lot of attention for the understanding of such flexible MOF materials. Such a dynamic structural behavior with proper host–guest interactions gives very interesting functions such as chemical separation, sensing and magnetic properties and so on.

1. Introduction

Metal–organic frameworks (MOFs) or porous coordination polymers (PCPs) have emerged as a rapidly evolving class of porous materials and commanded significant research attention in the materials and solid-state chemistry fields over the last two decades.¹ These porous crystalline solids are fabricated from bi- or multi-dentate organic linkers and held by metal-ion nodes in a periodic manner. The coordination bonded, self-assembled polymers are endowed with well-ordered porosity by virtue of a regularized arrangement of the building units in the infinite network. Because of the potential applications in the fields of storage, separation, sensing, and

ionic conduction, MOFs have been particularly sought after in recent years.^{2–5}

The key distinction of MOFs over congener porous materials is the class of third generation MOFs or soft porous crystals, which are related to the enzymic flexibility in such materials. Several studies over the last decade have demonstrated the advantage of examining molecular motions in crystalline polymeric solids.^{6–8} The relatively easy access to characterize nano-scale rearrangements unambiguously renders a remarkable additional advantage to these systems. Because of the assigned softness, MOFs have been able to function as guest-responsive host matrices without compromising on the polymeric character.^{9–12}

In pursuit of synthesizing pre-designed architectures in general and dynamic MOFs in particular, several donor groups have been investigated over the years. Among these groups, carboxylate and neutral N-heterocycle terminal ligands have been prominently preferred, with the former type dominating

Indian Institute of Science Education and Research (IISER), Pune. Dr. Homi Bhabha Road, Pashan, Pune, India 411008. E-mail: sghosh@iiserpune.ac.in;

Fax: +91 20 2589 8022; Tel: +91 20 2590 8076

[†]These authors have contributed equally.



Biplab Manna

Biplab Manna joined the Indian Institute of Science Education and Research (IISER), Pune in 2010 as a PhD student under the supervision of Dr Sujit K. Ghosh. He completed his BSc (Chemistry) from Calcutta University in 2008 and received an MSc from the Indian Institute of Technology, Roorkee in 2010. His research is mainly focused on developing ionic metal–organic frameworks for ion conduction and chemical sensing.



Aamod V. Desai

Aamod V. Desai obtained his BSc (Chemistry) from the University of Pune in 2011. Thereafter, he joined the Indian Institute of Science Education and Research (IISER), Pune, as an Integrated-PhD student. He is currently pursuing his doctoral research under the supervision of Dr Sujit K. Ghosh. His work is mainly focused on the investigation of N-donor ligand-based functional metal–organic frameworks for molecular recognition applications.



reports in the literature. Neutral N-donor linkers afford formation of the coordination bond leaving the maintenance of ionic equilibrium to the components of metal-salts used in the synthesis. This provides a simplistic route for the preparation of cationic frameworks where the counter-anion is left uncoordinated to the metal centre.^{13–19} Broadly, dynamic MOFs have relied on the flexibility possessed by the ligand and/or metal-ion with the reorganization primarily driven by desorption/resorption of neutral guest molecules, whereas N-donor ligand-based MOFs provide the additional aspect of enforcing an ionic stimuli (anion).

N-donor based ligands have been previously perceived as only pillars to increase the dimensionality in the carboxylate ligand linked frameworks.^{20,21} Research in recent years has shown the importance of such donor ligands in the domain of dynamic MOFs both in terms of understanding structural dynamism and probing structure–property correlation. This review focuses on the consolidation of the flexibility aspect in MOFs and the efficacy of N-donor ligands with representative examples concerning the structural aspects and the subsequent properties induced by the reorganization of the frameworks.

2. Scope of the review

The main focus of this particular review is to give an up-to-date report of flexible MOFs made of neutral N-donor ligands. A flexible MOF can be perturbed by various external stimuli (both chemical and physical) (Scheme 1).

The stimuli which render flexibility in N-donor ligand based MOFs are discussed in detail in this paper. In addition to this, different kinds of structure–property correlations are also thoroughly considered. It is worth noting that the neutral N-donor ligands are widely used classical ligands that can form MOFs upon binding to metal ions. Despite the availability of a few reviews on flexible MOF systems, the current

topic of concern lacks such reports in the literature. Several papers and reviews discussed in this article might attract attention for further investigations in such flexible MOF systems for various functions.

3. Synthesis routes to construct flexible frameworks

Generally, the uncontrolled polymeric self-assembly in MOFs makes it a challenging task to pre-design flexible systems. In principle, neutral N-donor ligand based MOFs are relatively well-suited for imparting structural dynamism because of the moderate strength of the primary bond involved. The free solvent molecules occluded during synthesis play the role of principal trigger to modulate the overall structure in MOFs. In particular, use of low boiling solvents facilitates structural dynamism because of the feasibility of the tendencies of such solvents to escape from the frameworks when keeping those MOF crystals away from mother liquor. Although anticipating coordination of guest molecules to metal sites is not possible, use of coordinating solvents and/or moderately interacting anions may afford a greater possibility to provide easily dissociable mono-dentate coordinated guests. The mandatory presence of the mono/multi-dentate anions, on some occasions uncoordinated to the metal centre, within N-donor ligand based MOFs offers the desirable stimulus of the exogenous anions to study solid-state structural transformations.²² In comparison physical stimuli driven investigations, light irradiated transformations in particular, have been found to be simpler in terms of design principles. With further work, the ambiguity in understanding the synthesis strategies of such systems can be significantly overcome.

4. Structural dynamism of N-donor ligand based MOFs

Neutral N-donor ligand based MOFs are generally observed to bear guest molecules which remain free or weakly coordinated to the metal centres. Variations of such guests inside the framework can cause alteration in framework structures because of the presence of relatively weaker coordinate bonds (M–N). Apart from such guest induced framework flexibility, dynamism can also be imparted *via* the perturbation of extra-framework anions.²³ Furthermore, inserting proper ligand functionality enhances the chance for molecular reorganization upon shining light (physical stimuli) on such frameworks. The single crystal to single crystal structural transformation method is one of the best direct and key techniques for understanding such molecular level chemical changes in such dynamic frameworks. Various exogenous stimuli (both chemical and physical) that induce structural dynamism will be discussed in detail in the subsequent sections of this paper.

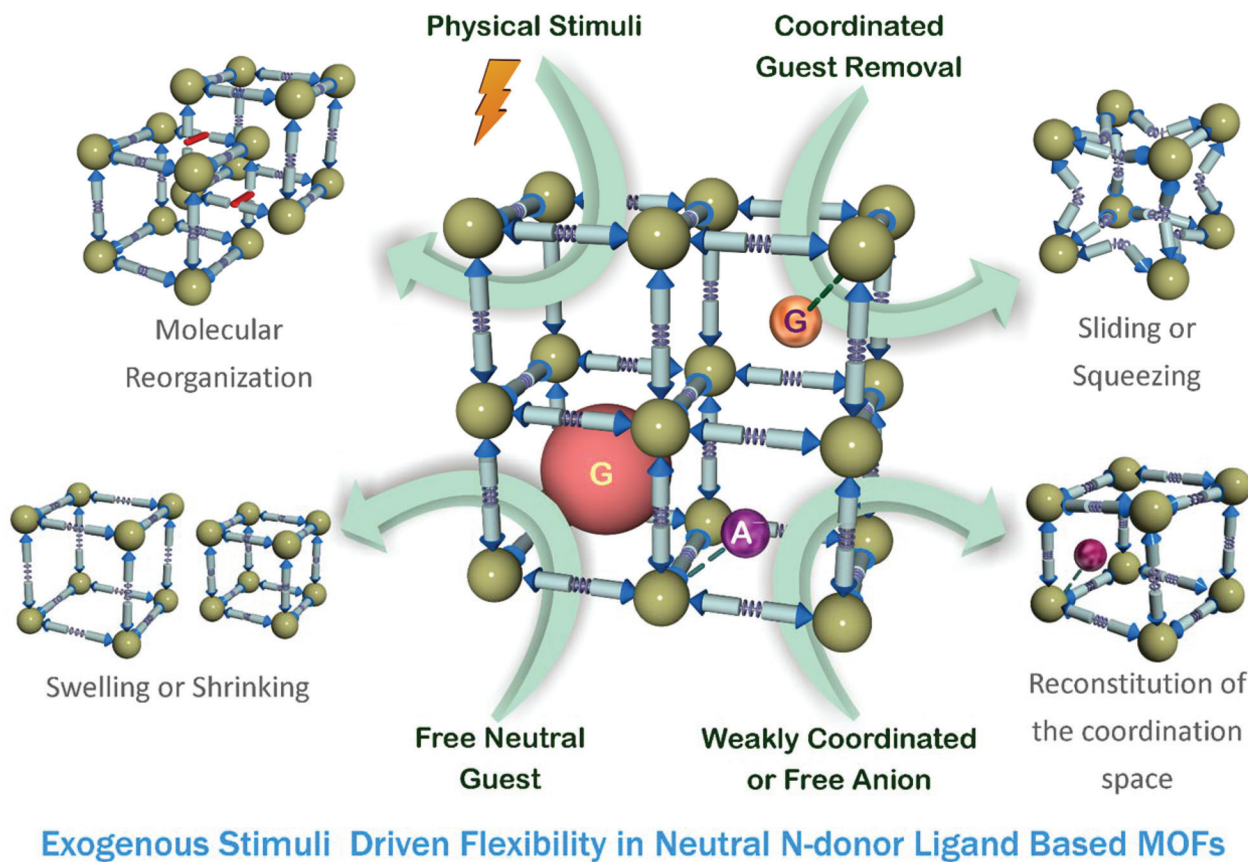


Sujit K. Ghosh

Dr Sujit K. Ghosh is an associate Professor in Chemistry at the Indian Institute of Science Education and Research (IISER), Pune, which he joined in 2009. He completed his PhD from the Indian Institute of Technology, Kanpur in 2006 under the supervision of Prof. Parimal K. Bharadwaj. He spent the following three years at Kyoto University, Japan as a JSPS and CREST postdoctoral fellow (host: Prof. Susumu Kitagawa). His

work focuses on the development and studies of functional metal–organic frameworks/porous coordination polymers suited for applications in the chemical industry, energy and environmental issues.





Scheme 1 Various exogenous stimuli (chemical and physical) that render flexibility in neutral N-donor ligand based MOFs.

4.1. Flexibility driven by chemical stimuli

Because of various exogenous stimuli (chemical) flexible MOFs undergo structural alteration. Such structural variations have often been validated well using single-crystal-to-single-crystal (SCSC) studies. Chemical stimuli (namely, counter anion, free guests, coordinated guests and gaseous guests) have been found to have a profound role in rendering flexibility in such MOFs.²⁴ In the following sections, the different types of chemical stimuli that cause flexibility in such MOF systems are categorized.

4.1.1. Anion induced framework flexibility. MOFs made of neutral N-donor ligands and metal ions possess counter anions in their framework lattice to make up the charge of the metal centers. Such counter anions that are present in the framework structure remain free in the voids or are weakly coordinated to metal ions and are found to interact with the cationic framework *via* various kinds of non-covalent interactions. Depending upon how strong these interactions are, replacement of such anions by incursive anions can be possible. It has been observed that such anion replacement process often makes a change in the structure of the original MOF system, thereby creating flexibility in the MOFs.^{25–27}

Yaghi *et al.* described long back structural transformation *via* reversible anion exchange.

For example, $[\text{Ag}(4,4'\text{-bpy})(\text{NO}_3)]_n$ (bpy = bipyridyl) transforms to an exchanged solid upon addition of an excess amount of potassium hexafluorophosphate.²⁸ On addition of excess potassium nitrate onto the exchanged solid, the original network is restored. Min and Suh showed anion dependent structural transformation in a cationic polynitrile-based network.²⁹ During the course of this anion exchange, one-dimensional (1D) $[\text{Ag}(\text{edtpn})(\text{NO}_3)]_n$ (edtpn = ethylenediamine-tetrapropionitrile) transforms to a two-dimensional (2D) layer structure $[\text{Ag}(\text{edtpn})(\text{CF}_3\text{SO}_3)]_n$, 2D box like network $[\text{Ag}(\text{edtpn})(\text{ClO}_4)]_n$ in a reversible and irreversible manner, respectively. $[\text{Ag}(\text{edtpn})(\text{ClO}_4)]_n$ can also be obtained from $[\text{Ag}(\text{edtpn})(\text{CF}_3\text{SO}_3)]_n$ in an irreversible manner (Fig. 1). The helical pitch of cationic infinite solids can be tuned by varying the extra framework anions. Jung *et al.* described a cationic coordination polymer (CP) $\{[\text{Ag}(\text{Py}_2\text{O})]\cdot\text{X}\}_n$ ($\text{X} = \text{NO}_3^-$, BF_4^- , ClO_4^- or PF_6^-) ($\text{Py}_2\text{O} = 3,3'\text{-oxybispyridine}$) as a smart helical spring³⁰ and this can stretch reversibly because of the counter anion exchange which can be correlated to the volume of anion (Fig. 2). Anion induced framework engineering has been shown in a 2D-interpenetrated cationic CP coordination polymer based on Cu(II) and 4,4'-bpy by Noro *et al.*³¹ The exchange of anion in an interpenetrated framework leads to an increase in the pore size of the overall framework.³² Replacement of free $\text{N}(\text{CN})_2^-$ in the framework $\{[\text{Ni}(\text{bpe})_2(\text{N}(\text{CN})_2)]-$



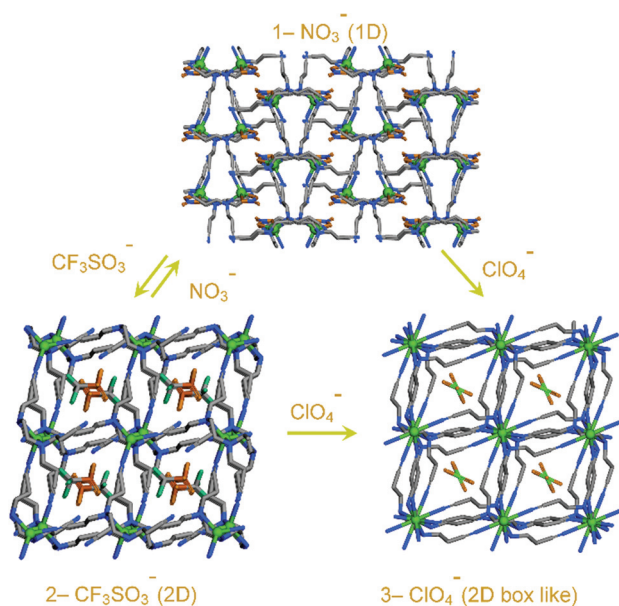


Fig. 1 Anion dependent structural transformation in a cationic polynitrile-based network.

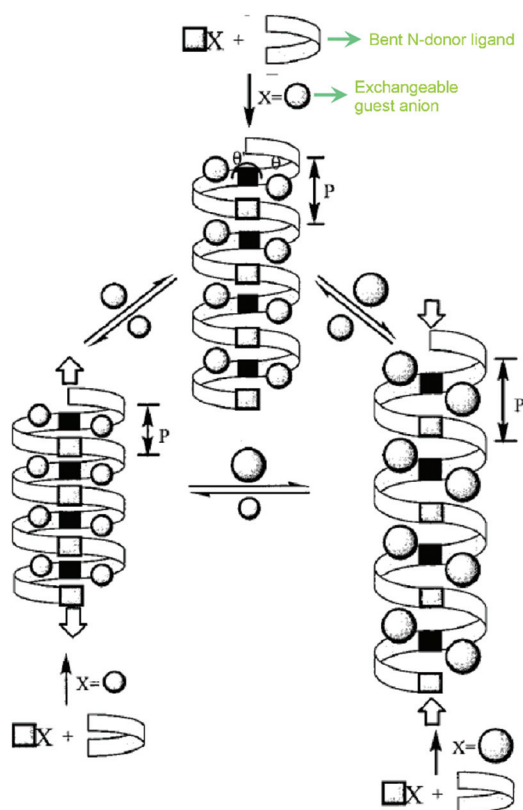


Fig. 2 Helical pitch tuning by anion exchange in a cationic solid. Reproduced with permission from ref. 30, Copyright (2000) American Chemical Society.

($\text{N}(\text{CN})_2$) $_n$ (bpe = 1,2-bis(4-pyridyl)ethane) by the smaller N_3^- causes mutual dislocations of the interpenetrating nets for enlargement in the effective porous area in the framework (Fig. 3). Such enhancement in the porous area in the framework affects the gas uptake property of the material. An N_3^- exchanged solid with a larger porous aperture shows a greater uptake of gas than the original framework. Our group devised a dynamic luminescent cationic framework $\{[\text{Zn}(\text{L}^1)-(\text{MeOH})_2](\text{NO}_3)_2 \cdot x\text{G}\}_n$ (L^1 = 4,4'-(ethane-1,2-diyl)bis(*N*-(pyridin-2-ylmethylene)aniline) which showed an interesting anion dependent structural dynamism depending on the size, shape and coordinating tendencies of various anions (NO_3^- , N_3^- , SCN^- , ClO_4^- and $\text{N}(\text{CN})_2^-$).³³ Li *et al.* showed acetate anion driven structural change in two isostructural three-dimensional (3D) cationic frameworks $\{[\text{M}(\text{L}^2)]_2(\text{ClO}_4)_2\}_n$ ($\text{M} = \text{Co}(\text{II})$ or $\text{Zn}(\text{II})$; L^2 = 4-(4-pyridyl)-3,5-bis(2-pyridyl)-1,2,4-triazole).³⁴ Hou *et al.* reported spherical halide anions (Cl^- and Br^-) dependent structural changes in a $\text{Cd}(\text{II})$ -based cationic framework $\text{Cd}(\text{L}^3)_2(\text{NO}_3)_2 \cdot 2\text{THF}$ (L^3 = 3,5-bis(3-pyridyl)-3-(3'-methylphenyl)-1,3,4-oxadiazole).³⁵

In an early work, our group devised two MOF-based supramolecular isomers $\{[\text{Cd}(\text{L}^4)_2(\text{ClO}_4)_2] \cdot \text{toluene} \cdot \text{MeOH}\}_n$ and $\{[\text{Cd}_2(\text{L}^4)_4(\text{ClO}_4)_4] \cdot 3(\text{mesitylene})\}_n$ (L^4 = 1,4-bis(4-pyridyl)-2,3-diaza-1,3-butadiene) which show anion induced structural changes for incoming strongly coordinating N_3^- and SCN^- anions because of the replacement of weakly coordinated ClO_4^- anions in the original frameworks.³⁶ Anion driven structural transformation has also been exploited for visual colorimetric detection of anions.

For example, a 1D amide functionalized $\text{Cu}(\text{II})$ -based framework $\{[\text{CuL}^5_2(\text{NO}_3)_2] \cdot o\text{-xylene} \cdot \text{DMF}\}_n$ changes its structure to a 2D framework $\{[\text{CuL}^5_2(\text{SCN})_2] \cdot x\text{G}\}_n$ (L^5 = 5-*tert*-butyl-N1,N3-di-(pyridin-4-yl)isophthalamide) upon treatment with SCN^- and this is accompanied by a color change from blue to green which can be seen by the naked eye (Fig. 4).³⁷

4.1.2. Alteration of free guests. MOFs composed of such neutral N-donor ligands are often found to have free guests

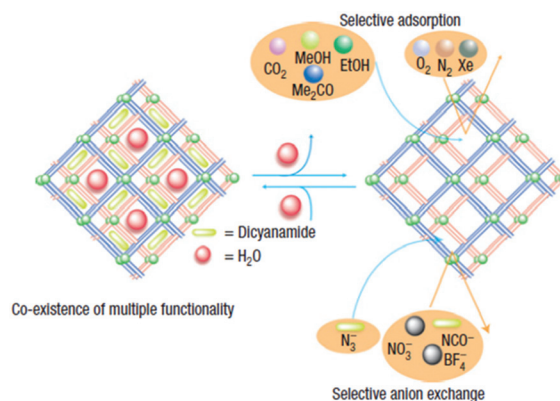


Fig. 3 Schematic illustration of mutual sliding of nets induced by anion exchange in a cationic MOF. Reprinted with permission from ref. 32, Copyright (2007) Macmillan Publishers Ltd.



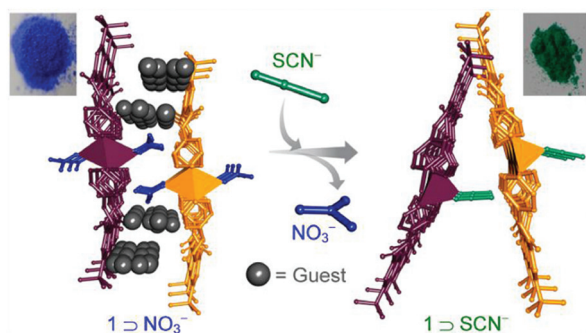


Fig. 4 Visual colorimetric anion sensing in a dynamic cationic MOF. Reproduced with permission from ref. 37, Copyright (2015) Wiley-VCH.

inside the framework. Those free guests are bound to the network by weak molecular interactions and thereby enhance the chances for the variations of such guests inside the framework. Alteration of such guests within the framework leads to structural changes of the system. Because of this structural flexibility, third generation MOFs possess bistable or multi-stable states and show enhanced host-guest interactions compared to rigid MOFs.³⁸

Carlucci *et al.* showed sponge like behavior in a Cu(II)-based framework $[\text{Cu}_5(\text{bpp})_8(\text{SO}_4)_4(\text{EtOH})(\text{H}_2\text{O})_5](\text{SO}_4) \cdot \text{EtOH} \cdot 25.5\text{H}_2\text{O}$. (bpp = 1,3-bis(4-pyridyl)propane) The compound exhibits guest dependent reversible structural transformation towards desolvation and resolution.³⁹ Noro *et al.* reported a 3D framework $\{[\text{Cu}(\text{AF}_6)(4,4'\text{-bpy})_2] \cdot x\text{H}_2\text{O}\}_n$ that changes to a 2D interpenetrating network $\{[\text{Cu}(4,4'\text{-bpy})_2(\text{H}_2\text{O})_2] \cdot \text{AF}_6\}_n$ (A = Si, Ge or Ti) in an aqueous environment.³¹ Guest driven sliding of 2D layers in a MOF has been demonstrated by Biradha *et al.*⁴⁰ Upon guest exchange, the 2D framework $[\text{Ni}(\text{L}^6)_2(\text{NO}_3)_2 \cdot 4(o\text{-xylene})]_n$ transformed to $[\text{Ni}(\text{L}^6)_2(\text{NO}_3)_2 \cdot 1.7(\text{mesitylene})]_n$ (L^6 = 2,4,6-tris(4-pyridyl)triazine) with a significant increase in the channel dimensions (Fig. 5).

In continuation, Biradha *et al.* also exhibited a dynamic MOF which remarkably shrinks or swells upon desorption or adsorption of the guest molecule.⁴¹ When the crystals of the framework $\{[\text{ZnI}_2]_3(\text{L}^6)_2 \cdot 6\text{C}_6\text{H}_5\text{NO}_2\}_n$ were kept in the open atmosphere for one day, they changed to a new phase with a remarkable compression of the network without losing their

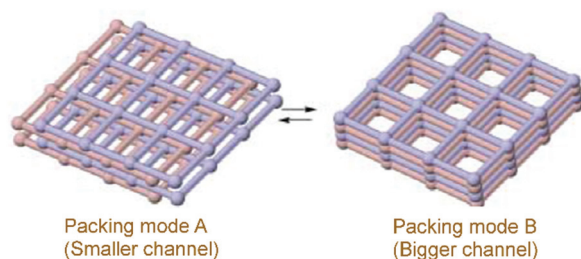


Fig. 5 Guest induced sliding of 2D layers in a reversible way. Reproduced with permission from ref. 40, Copyright (2002) Wiley-VCH.

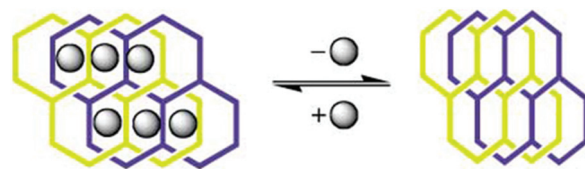


Fig. 6 Schematic representation of guest dependent shrinking or swelling in a dynamic MOF. Reproduced with permission from ref. 41, Copyright (2002) Wiley-VCH.

single crystalline nature. Interestingly, upon immersing the crystals of the new phase into nitrobenzene, the parent crystals are restored (Fig. 6). A reversible crystal structure to amorphous structural transformation has been examined in an amide functionalized 3D MOF by Uemura *et al.* The 3D framework $\{[\text{Co}(\text{NCS})_2(4\text{-peia})_2] \cdot 4\text{Me}_2\text{CO}\}_n$ (4-peia = *N*-(2-pyridin-4-yl-ethyl)-isonicotinamide) transformed to an amorphous state upon complete removal of the guest acetone molecules and changed back to their original state when the amorphous phase was exposed to acetone vapour.⁴² Guest dependent structural flexibility has also been shown in a cage-based coordination network.⁴³ For example $[\text{Co}_3(\text{SCN})_6(\text{TPT})_4]_n$ transformed to $[\text{Co}_3(\text{SCN})_6(\text{TPT})_4(\text{TTF})]_n$ (TPT = 2,4,6-tris(4-pyridyl)-triazine) while taking up guests (tetrathiafulvalene) with a change of the crystal system from cubic to tetragonal (Fig. 7).

Recently, Bloch and Sumbly demonstrated guest induced crystal to crystal expansion and contraction in a Ag(I)-based 3D MOF $\{[\text{Ag}(\text{dpzm})]\text{ClO}_4 \cdot 1.1(\text{DMSO}) \cdot 0.9(\text{ETOAC})\}_n$ (dpzm = di-2-pyrazinylmethane)(1 As-made).⁴⁴

Upon guest exchange with dichloromethane (DCM), 1 As-made changed to 1 DCM in a crystal to crystal manner accompanied by a dramatic contraction of 3.09 Å along the *c* axis. But 1 DCM expanded to 1 acetone in a reversible fashion after soaking in acetone for few days (Fig. 8). In continuation, the authors reported a similar series of MOFs which exhibited similar breathing transformations upon guest exchange.⁴⁵ In a very recent report our group showed a 3D interpenetrated cationic MOF $\{[\text{Zn}(\text{L}^7)_2(\text{NO}_3)_2 \cdot x\text{G}]\}_n$ (L^7 = [*E*]-*N*'-[1-(pyridin-4-yl)ethyldene]hydrazine carbohydrazide] which upon dynamic structural transformation transformed to a 2D MOF (Fig. 9).⁴⁶

In continuation, guest driven inherent dynamic behaviour of a Cd(II)-based cationic MOF has been shown in research by us.⁴⁷ Upon air-drying, the as-made MOF $\{[\text{Cd}(\text{L}^4)_3 \cdot (\text{ClO}_4)_2] \cdot x\text{G}\}_n$ transforms to a 2D MOF $\{[\text{Cd}(\text{L}^4)_2(\text{OH}_2)_2(\text{ClO}_4)_2] \cdot \text{THF}\}_n$ via the loss of a coordinated ligand molecule (Fig. 10).

4.1.3. Flexibility driven by variation of coordinated guests.

An easy route to achieve perturbation in the coordination sphere of metal nodes can be by substitution of coordinated neutral molecules. By virtue of the mono-denticity of such coordinated molecules, they can easily be substituted or the framework can reorganize itself to make up for the presence of open metal sites. In certain cases the regulation of flexibility by coordinated guests has been associated with change in



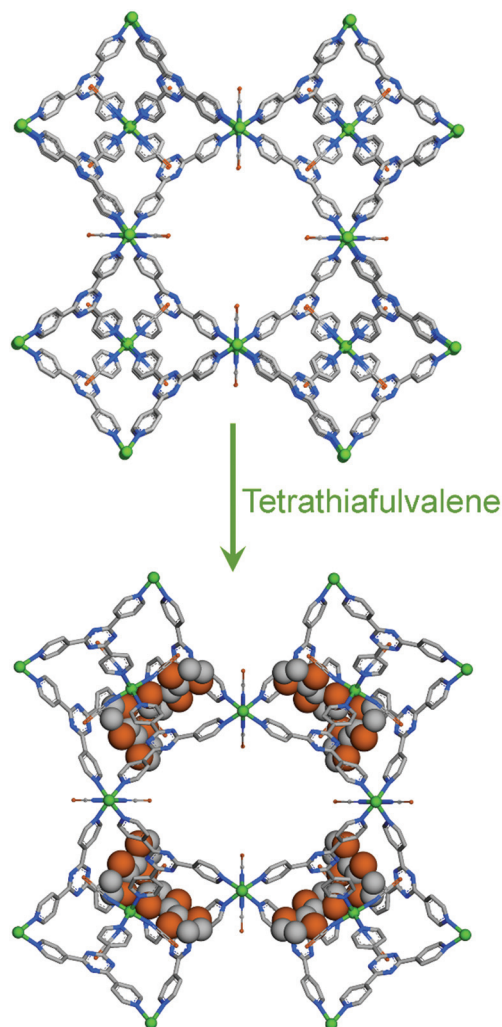


Fig. 7 TTF mediated structural changes in a networked molecular cage.

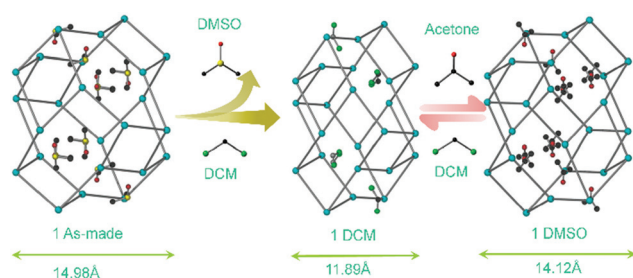


Fig. 8 Guest mediated crystal expansion and contraction in a cationic MOF. Reproduced from ref. 44, Copyright (2012) The Royal Society of Chemistry.

certain physical properties.^{48,49} Some of the key results in recent years are presented next.

Bradshaw *et al.* demonstrated the flexibility of a MOF by regulating the coordinated neutral guest *via* thermal stimuli.⁵⁰ The authors synthesized a neutral compound $[\text{Co}_2(\text{bipy})_3]^-$

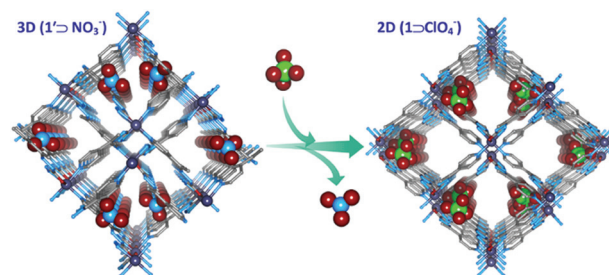


Fig. 9 Anion induced structural transformation from a 3D framework to 2D sheets. Reproduced with permission from ref. 46, Copyright (2014) American Chemical Society.

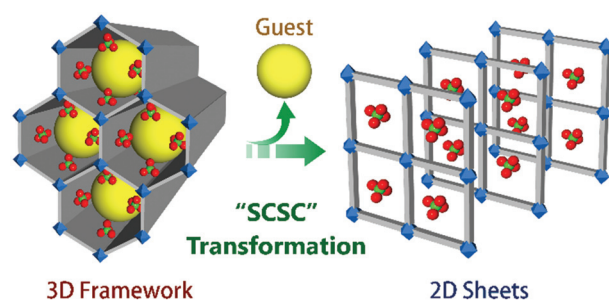


Fig. 10 SCSC transformation from a 3D porous framework to 2D non porous sheets. Reproduced from ref. 47, Copyright (2015) The Royal Society of Chemistry.

$(\text{SO}_4)_2(\text{H}_2\text{O})_2][\text{bipy}](\text{CH}_3\text{OH})$, (bipy-4,4'-bipyridyl) bearing an uncoordinated ligand and a solvent molecule. Upon heating the coordinated water molecules were replaced by the uncoordinated species in the parent compound. This change was accompanied by alteration in the structure of the compound to give a new phase. The reversibility of this experiment and the hypothesis was validated by controlled cooling in the presence and absence of atmospheric moisture (Fig. 11).

In a similar report, Aslani and Morsali showed that drastic structural changes of a coordination polymer occurred by removal of a coordinated solvent upon heating.⁵¹ A 1D coordi-

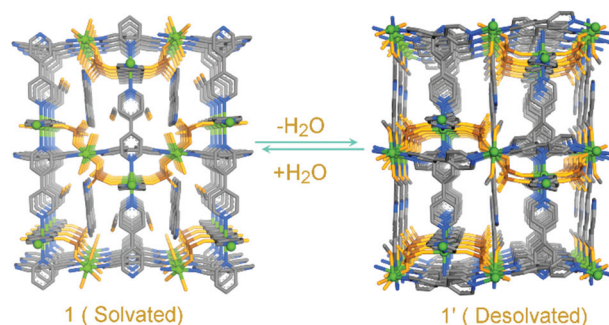


Fig. 11 Rearrangement of the coordination environment by loss of bound water.



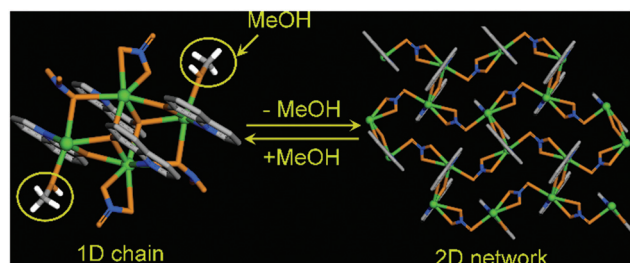


Fig. 12 Increase in framework dimensionality via loss of coordinated MeOH.

nation polymer $[\text{Pb}_2(8\text{-Quin})_2(\text{NO}_3)_2(\text{MeOH})]$, (8-Quin = 8-hydroxyquinoline) was synthesized which was converted to a 2D structure $[\text{Pb}(8\text{-Quin})-(\text{NO}_3)]$ when the compound was heated at 165–170 °C. The Pb–O bond with MeOH was replaced by a Pb–O bond from the NO_3 anion. This transition was also accompanied by a visual change in the colour of the crystal.

The resolution of the compound to its parent phase was achieved when this was dipped in MeOH and the authors observed such multiple reversible transformations (Fig. 12). Zhuang *et al.* extended this concept through fabrication of a Cd(II) based MOF $[(\text{Cd}(\text{ImBNN})_2(\text{CF}_3\text{SO}_3)_2)\cdot\text{guest}]_n$, (ImBNN = 2,5-bis[4'-(imidazol-1-yl)phenyl]-3,4-diaza-2,4-hexadiene) bearing weakly coordinated triflate anions and free toluene molecules.⁵² Upon removal of the toluene molecules by heating, the metal–anion bond is weakened. When the compound was cooled under atmospheric conditions, one water molecule replaced the triflate anion leading to slight structural changes.

The original phase of the compound was restored by heating the water coordinated phase in toluene at 110 °C. Toluene molecules occupied the free space of the MOF and both the anions were coordinated to the metal centre. This report yields an example of neutral guest molecules competing with charged anions for metal-coordination leading to structural changes (Fig. 13).

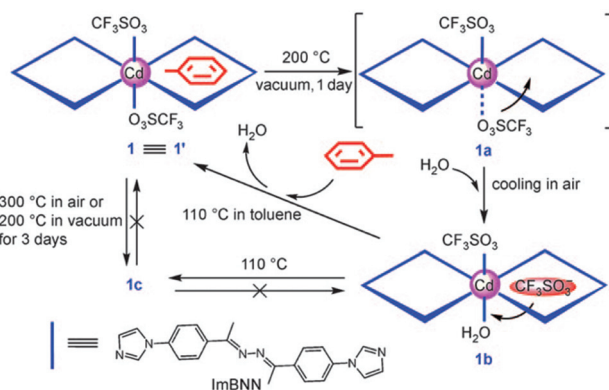


Fig. 13 Schematic illustration of structural transformation upon loss and gain of coordinated guests. Reproduced with permission from ref. 52, Copyright (2009) Wiley-VCH.

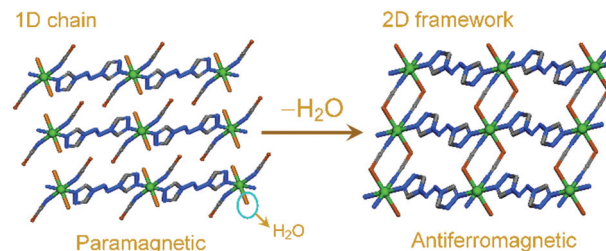


Fig. 14 SCSC transformation upon dehydration accompanied by changes in magnetic properties.

Liu *et al.* observed the structural changes which accompanied the promising magnetic regulation in a MOF-based material.⁵³ The authors synthesized a Co(II) centred 1D coordination polymer $[\text{Co}(\text{SCN})_2(\text{bta})(\text{H}_2\text{O})_2]$, (bta = 1,2-di(4*H*-1,2,4-triazol-4-yl)diazene) comprising water molecules in the coordination sphere of the metal ion. Upon desolvation at 100 °C the authors observed structural changes together with an increase in dimensionality of the packing. Interestingly, a magnetic variation was noticed during this transition from the simple paramagnetic behaviour of the parent compound to the antiferromagnetic nature of the desolvated phase (Fig. 14). As discussed in the previous section, a MOF exhibiting anion and guest dependent structural changes has been reported.³³ The as-synthesized compound on air-drying underwent noticeable structural modifications, *via* SCSC characterization, and accompanied with a change in the coordination environment of the metal centre. The coordinated methanol molecules were replaced by H_2O from atmospheric moisture.

The affinity of the framework towards several hydrophilic adsorbates was studied further, and this yielded size selective uptake (Fig. 15).

4.1.4. Flexibility driven by gas adsorption. Another route to flexibility is the structural changes effected by the interaction of gas molecules with the framework components under pressure.^{54,55} Because of the precise understanding of gate opening phenomenon in MOFs, few N-donor ligand based MOFs have been found to follow such behaviour. Some

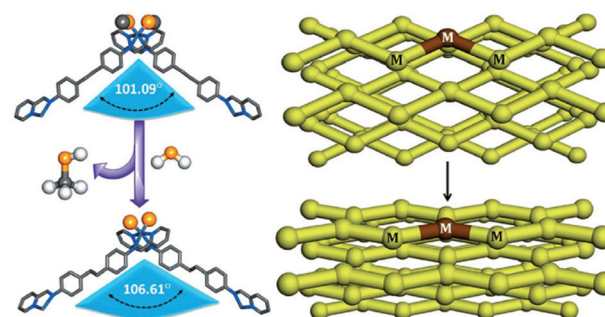


Fig. 15 Guest triggered dynamic structural transformation in a flexible cationic MOF. Reproduced with permission from ref. 33, Copyright (2013) Wiley-VCH.



of the important recent examples have been discussed in this review.

In an early work Li and Kaneko studied the adsorption behaviour for a Cu(II) MOF $[\text{Cu}(\text{bpy})(\text{BF}_4)_2(\text{H}_2\text{O})_2 \cdot (\text{bpy})]_n$.⁵⁶ A similar response to different gases such as N_2 , Ar, and CO_2 was observed in which the relatively non-porous packed compound was able to adsorb a significant amount of these gases after tuning of the gate pressures. The authors attribute the breach of hydrogen bond assisted packing of the compound (because of the presence of BF_4^- anions) to the respective gate pressures of the adsorbed gases (Fig. 16).

Kondo *et al.* reported a dual flexible MOF system which underwent structural alterations upon guest removal and later upon adsorption of gas molecules.⁵⁷ The authors synthesized a 3D MOF $\{[\text{Cu}(\text{bpy})(\text{BF}_4)_2(\text{H}_2\text{O})_2] \cdot (\text{bpy})\}$, which underwent a dimensionality reduction to form a 2D framework having the formula $[\text{Cu}(\text{BF}_4)_2(\text{bpy})_2]$ (Fig. 17).

Upon studying CO_2 adsorption for the 2D compound, a notable change in the interactions between two 2D sheets was found to happen, which was ascribed to the clathrate formation between the gaseous molecules and the framework in the adsorbed state.

In a continuation of the research, Kondo *et al.* demonstrated the flexibility of dynamic frameworks upon inclusion of

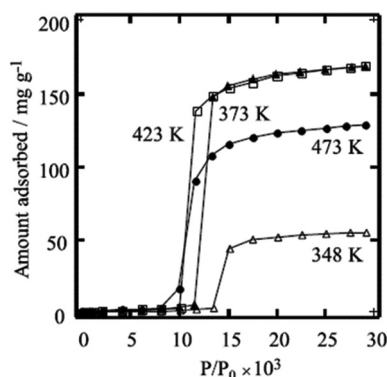


Fig. 16 Gate opening behaviour for CO_2 adsorption in a Cu(II) MOF. Reprinted with permission from ref. 56, Copyright (2001) Elsevier.

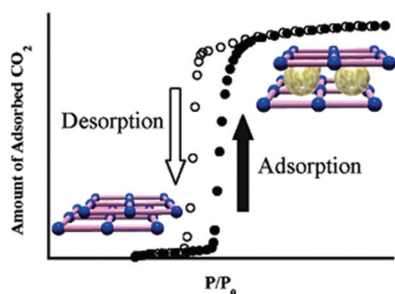


Fig. 17 Schematic representation of clathrate formation during CO_2 adsorption. Reproduced with permission from ref. 57, Copyright (2006) American Chemical Society.

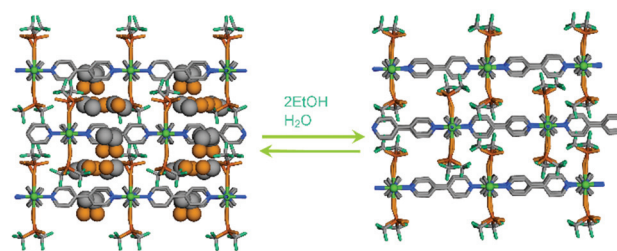


Fig. 18 Guest induced reversible structural transformation in a flexible MOF.

gaseous molecules by constructing a Cu(II) based 2D MOF $[\text{Cu}(\text{bpy})_2(\text{OTf})_2] \cdot 2\text{EtOH} \cdot \text{H}_2\text{O}$.⁵⁸ An isostructural MOF with MeOH occluded guests has also been used in this work to study the effect of free guest molecules. Upon desolvation the parent compound was subjected to low temperature N_2 adsorption, where the authors observed the typical gate-opening behaviour at $P/P_0 = 0.16\text{--}0.22$. Powder X-ray diffraction (PXRD) measurements at the first step of gate-opening showed the structural changes corresponding to the expansion of the 2D layers in the compound (Fig. 18).

Kondo *et al.* expanded their previous work by fabricating a series of MOFs having the same composition, namely, $[\text{Cu}(\text{bpy})_2(\text{OTf})_2]_n$, in different solvent combinations.⁵⁹ In the 2D architectures the authors observed stepwise adsorption for gases such as N_2 , CO_2 , Ar, CH_4 . This behaviour was not observed for the 3D compounds where a Type-I adsorption profile was seen in all cases and the uptake followed a molecular sieving effect. The layered 2D structures have been thought to undergo sliding to accommodate the respective adsorbents (Fig. 19).

In a similar work, Kotani *et al.* synthesized a Cu(II) based 1D coordination polymer, namely, $\text{Cu}(\text{bpp})_2(\text{BF}_4)_2$, which underwent structural changes upon desolvation.⁶⁰ Low temperature N_2 adsorption was performed on the guest-free phase, which did not show any uptake. When CO_2 adsorption was performed on the same phase under moderate conditions, the compound exhibited a gate-opening adsorption profile.

4.2. Flexibility arising from physical stimuli

MOFs are known to be responsive to physical stimuli because of the dual advantages bestowed by the metal ions and organic ligands. Among them, light irradiated structural changes have commanded significant attention among other physical stimuli. Park *et al.* have pioneered the domain of cycloaddition reaction based structural changes in olefin-bearing ligands.⁶¹ Because of the lack of chemical stability in neutral N-donor ligand based MOFs, solid-state reactions are more favoured for investigation of structural subtleties. The prominent reports among such class of studies are discussed next.

In an early study on this aspect, Nagarathinam and Vittal reported anisotropic movements within a CP structure.⁶² The authors synthesized a CP $[\text{Ag}(\mu\text{-bpe})(\text{H}_2\text{O})](\text{CF}_3\text{CO}_2) \cdot \text{CH}_3\text{CN}$, (bpe-4,4'-bipyridylethylene) which upon desolvation was



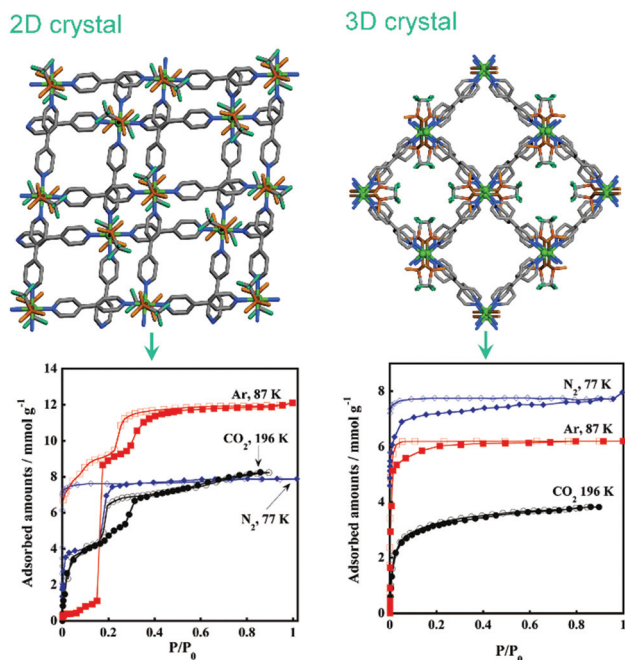


Fig. 19 Differential adsorption behaviours in two MOFs having the same composition. Reproduced with permission from ref. 59, Copyright (2011) American Chemical Society.

subjected to photo-irradiation. The resultant compound was found to bear the desired cyclobutane rings, as confirmed using nuclear magnetic resonance (NMR) studies. Although crystallographic evidence could not be provided for the desolvated structure, the authors crystallized a similar compound bearing coordinated anions $[(\mu\text{-O}_2\text{CCF}_3)_2\text{Ag}]_2(\mu\text{-bpe})_2 \cdot \text{H}_2\text{O}$ which supported their claim of the formation of a ladder-like CP. The authors hypothesized the reorganization of the compound after desolvation to align the ethylene groups making them suitable for a cycloaddition reaction.

Using this knowledge, Nagarathinam and Vittal used a photochemical cycloaddition reaction to show the formation of the intermediate upon dehydration of a 1D coordination polymer $[\text{Cd}(\text{bpe})(\text{CH}_3\text{COO})_2(\text{H}_2\text{O})]_n$.⁶³ The synthesized compound upon dehydration was subjected to ultraviolet (UV) irradiation to achieve 100% formation of the cyclobutane ring. Compared to this result, only 33% conversion was noted when the parent compound was irradiated. Thus, among the two possible intermediates, the authors could conclusively infer the formation of the 1D ladder-like structure formation upon dehydration.

In continuation, Peedikakkal and Vittal synthesized a triple-stranded ladder CP $[\text{Pb}_3(\text{bpe})_3(\text{O}_2\text{CCF}_3)_4(\text{O}_2\text{CCH}_3)_2]_n$, which underwent a solid state cycloaddition reaction.⁶⁴ A two-step reaction was observed which occurred because of the anisotropic movements among the adjacent strands of the CP. In the first step ~67% photodimerization was noted which corresponded to the reaction between a pair of bpe ligands with a triple-stranded structure. The second step involved the remain-

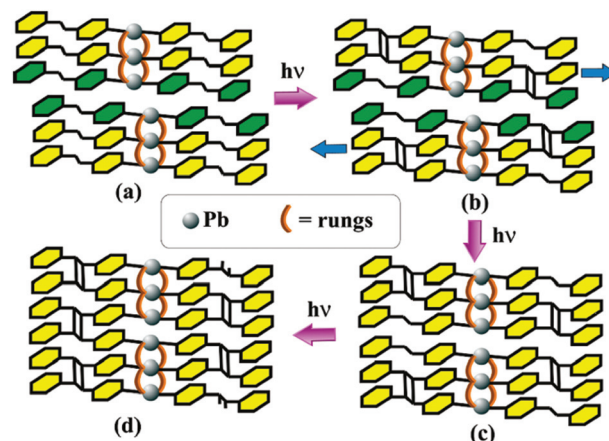


Fig. 20 Schematic representation of a two-step photo induced dimerization. Reproduced with permission from ref. 64, Copyright (2010) American Chemical Society.

ing 33% reaction with cooperative movements in the polymeric structure to align the unreacted olefin for completing the dimerization (Fig. 20).

In a parallel manner, the authors have reported development of CPs using the photoactive monomers. By employing a Zn(II) salt and a mono-dentate N-donor ligand, a monomer $[\text{ZnBr}_2(4\text{spy})_2]$, (spy = *trans*-2-fluoro-4'-styrylpyridine) was synthesized.⁶⁵ Upon irradiation, a dimer was obtained which could be reversibly brought back to the monomer by heating. When the dimer was irradiated for a longer duration, a 1D CP was synthesized, the reaction of which was not found to be reversible. The authors further observed a drastic photoluminescence change together with the photopolymerization (Fig. 21). More recently, Medishetty *et al.* undertook a comprehensive study of the photo-activeness of monomeric systems bearing photo-active olefinic groups.⁶⁶

A series of Ag(I) centred compounds were synthesized by varying the ligand and anion. All the compounds were found to undergo photo-induced cycloaddition reactions, which provided conclusive evidence for the effect of the shape, nature, and size of the counter-anions on the photo-activity of the compound.

5. Flexibility driven physical and chemical properties

The flexibility endowed by dynamic MOF systems permits investigation of these materials for controlled regulation of certain physical or chemical properties and/or determining their suitability for prospective real-time applications.^{67,68} A few important results relating to this are discussed next.

In an early report, Halder *et al.* found a flexible Fe(II) centred MOF which exhibited guest-dependent spin-crossover (SCO) properties.⁶⁹ The authors synthesized a MOF $\text{Fe}_2(\text{azpy})(\text{NCS})_4(\text{EtOH})$, (azpy = *trans*-4,4'-azopyridine) where a weak



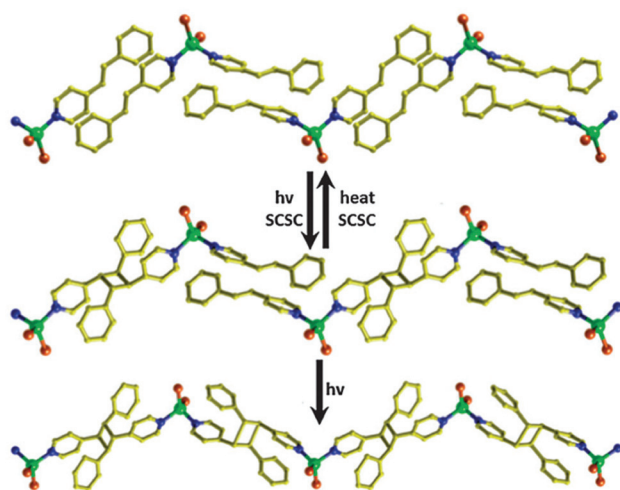


Fig. 21 Photo induced cycloaddition reaction in the formation of 1D CP. Reproduced from ref. 65, Copyright (2013) The Royal Society of Chemistry.

interaction between O–H...S was found to prevail. Temperature dependent magnetic susceptibility was recorded followed by desolvation. In the guest-free form, the slightly bent thiocyanate ions were found to be aligned in a linear fashion, which gave rise to structural changes.

Upon measurement of magnetic susceptibility for this phase of the compound, a significant change in the crossover onset temperature was observed. To corroborate with the proposed mechanism, the desolvated phase was dipped separately in EtOH, MeOH and PrOH to obtain resolvated guest included compounds. These phases were found to show similar susceptibility patterns (Fig. 22). In continuation and carrying on from the previous work, the authors synthesized a Fe(II) MOF $[\text{Fe}(\text{NCS})_2(\text{bpbd})_2] \cdot \{\text{acetone}\}$, (bpbd = 2,3-bis(4'-pyridyl)-2,3-butanediol) comprising free hydroxyl bearing groups in the ligand.⁷⁰ The O–H...S interactions, as previously noted, were found to persist in the interpenetrated packing of the compound, rendering inflexibility to the compound and keeping out any host–guest non-covalent interaction.

Temperature dependent, subtle structural changes were observed for the compound. Temperature dependent magnetic susceptibility measurements were carried out for the as-synthesized phase and the desolvated phase, and these displayed an almost similar tendency, reaffirming the hypothesis of employing a free hydroxyl-based ligand (Fig. 23).

Investigating further, Neville *et al.* have reported a few studies including SCSC transformations of Fe(II) SCO frameworks,⁷¹ a mechanistic understanding of multi-step SCO transitions⁷² and structural modulations by guest-dependent SCO behaviour of seemingly iso-structural MOFs.⁷³

In a distinct domain of investigation, Maji *et al.* demonstrated the use of the dynamic nature of a two-fold interpenetrated 3D MOF $\{[\text{Ni}(\text{bpe})_2(\text{N}(\text{CN})_2)](\text{N}(\text{CN})_2(5\text{H}_2\text{O}))\}_n$, for controlled gas adsorption.³² The compound bears two kinds of porous channels occupied by the solvent (water) and anions

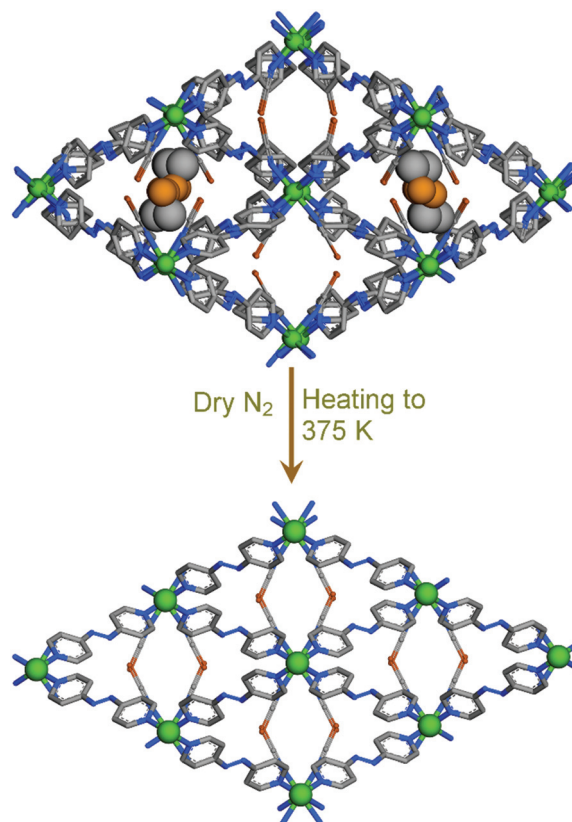


Fig. 22 Guest dependent spin crossover behaviour in dynamic MOF.

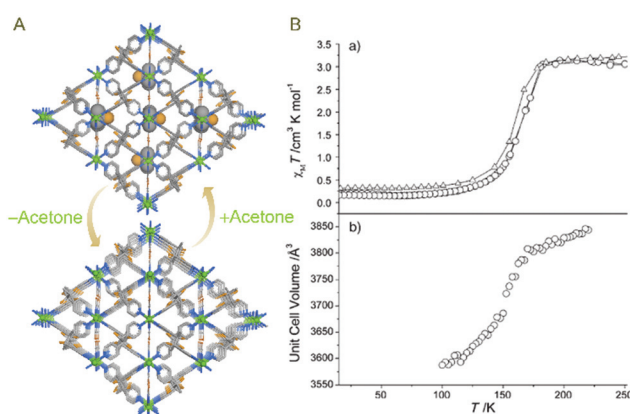


Fig. 23 (A) Guest induced reversible structural changes (B) without affecting the magnetic behaviour of host framework. Reproduced with permission from ref. 70, Copyright (2007) Wiley-VCH.

separately. Upon desolvation, a noticeable change in colour was observed which was restored upon resolution. The dehydrated framework bestows a dual function to the compound where N_3^- anions are exchanged selectively with $\text{N}(\text{CN})_2^-$ ions in competition with NCO^- , NO_3^- , BF_4^- . The anion-exchanged phase in turn renders an additional feature of significantly enhanced gas adsorption by enlarging the window of the



larger pore and squeezed smaller pore, by virtue of the linear shape of the exchanged anion.

In a parallel work, Noro *et al.* have reported selective CO₂ adsorption by a 2D PCP.⁷⁴ The authors synthesized a Cu(II) centred MOF $\{[\text{Cu}(\text{PF}_6)(4,4'\text{-bpy})_2(\text{MeOH})]\cdot\text{PF}_6\cdot 3\text{MeOH}\}_n$, which upon desolvation underwent a change in structure to $[\text{Cu}(\text{PF}_6)_2(4,4'\text{-bpy})_2]_n$ where the uncoordinated PF₆[−] ions replaced the coordinated MeOH molecules in the parent compound. Upon checking the adsorption behaviour towards various gases such as CO₂, N₂, O₂ and Ar, a strong preference and high uptake for CO₂ at both low and high temperatures was observed. Theoretical calculations revealed the affinity of the fluorine ions towards CO₂ as the primary driving force for the high uptake and low energy regeneration (Fig. 24). CO₂ separation based on a similar adsorption mechanism has also been studied by Nugent *et al.* using a SiF₆^{2−} bridged N-donor ligand based MOFs.⁷⁵

Apart from these properties, such dynamic MOFs have found use as luminescent probes for a few important neutral and ionic analytes. Wang *et al.* have reported a comprehensive study of structure–property correlation in a flexible MOF $[\text{Cu}(\text{CN})_3\text{L}_8(\text{H}_2\text{O})(\text{CH}_3\text{CH}_2\text{OH})]_n$, (L₈ = 2,6-bis((3,5-dimethyl-1H-pyrazol-4-yl)methyl)pyridine) acting as a solvent-vapour sensor for use with the naked eye.⁷⁶ The compound has an interpenetrated structure which leaves room for pre-organization towards the guest according to its nature and size. Upon desolvation a change in PXRD was found which was restored upon resolution. Additionally this process was visible under UV irradiation with a notable difference of emitted light frequency.

A drastic change in the luminescence of the desolvated compound was observed when it was immersed in several sol-

vents such as acetonitrile, benzene, cyclohexane, cyclohexene, dichloromethane, ethanol, methanol, and tetrahydrofuran. To gain molecular insights into these photophysical observations, the authors synthesized a few compounds which had the guest molecules included. Furthermore, to understand the real-time applicability of the compound, solid-state fluorescence studies towards acetonitrile vapours were investigated. Together with the response towards different concentrations of acetonitrile vapours, the change in fluorescence for solvent mixtures was also examined (Fig. 25).

As briefly discussed in the earlier sections,³³ a dynamic, luminescent PCP exhibiting guest and anion dependent flexibility has been reported. The ligand used is a π -conjugated molecule, which bestows ligand-based fluorescence to the compound. The parent compound was found to exhibit high solid-state emission primarily because of the intraligand charge transfer. Upon anion-exchange, the fluorescence intensities and positions were perturbed because of plausible changes in the interactions of the anions with the framework (Fig. 26). Such a visible, non-sophisticated tool to monitor

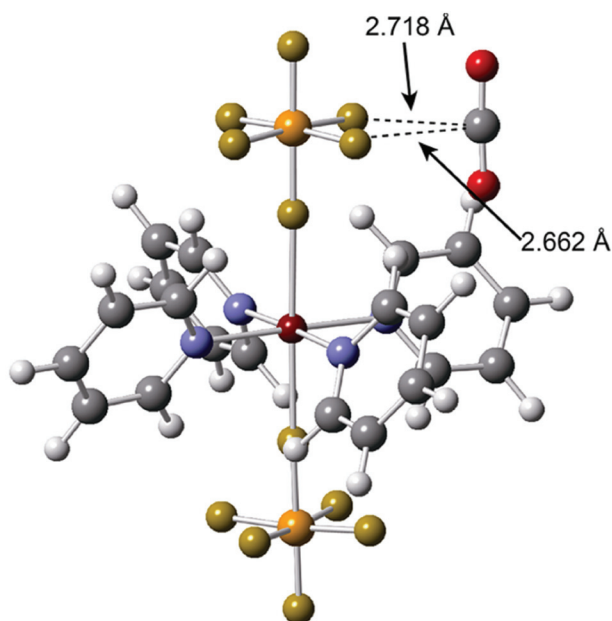


Fig. 24 Optimized structure of $[\text{Cu}(\text{PF}_6)_2(\text{pyridine})_4]$ with CO₂ obtained using density functional theory. Reproduced with permission from ref. 74, Copyright (2013) American Chemical Society.

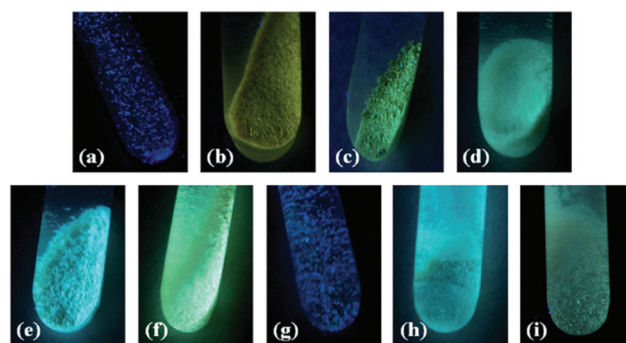


Fig. 25 Naked eye fluorescence changes in a guest responsive dynamic MOF upon immersion in several volatile organic solvents (a) parent compound, (b) acetonitrile, (c) methanol, (d) ethanol, (e) tetrahydrofuran, (f) benzene, (g) cyclohexane, (h) cyclohexane (longer time), (i) dichloromethane. Reproduced from ref. 76, Copyright (2013) The Royal Society of Chemistry.

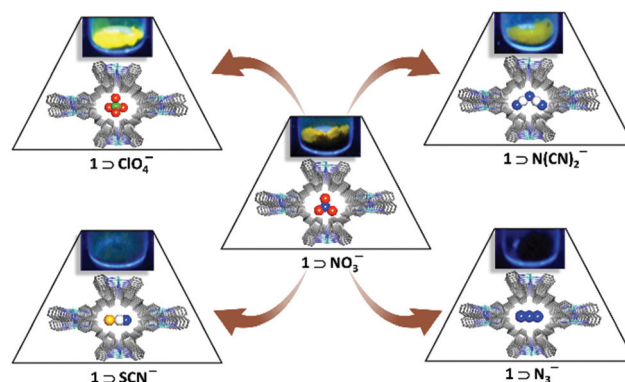


Fig. 26 Anion responsive tunable luminescent behaviour in a dynamic cationic MOF.



anion-exchange/capture process may be useful for important applications in the material/biological fields.

6. Conclusions and future prospects

In this review, a systematic update in the field of flexible MOFs composed of neutral N-donor ligands has been given. A range of exogenous stimuli which trigger structural flexibility/softness to such MOF systems have been discussed. Various chemical stimuli such as framework anions, free guests, co-ordinated guests and gaseous guests together with physical stimuli (light) have been used in creating flexibility in these kinds of MOF systems. Various flexibility driven functions in these materials (e.g., separation of gases, sensing of small molecules, magnetic properties and so on) which include proper host-guest interactions have been thoroughly examined. For a better understanding of such a flexible MOF system, it is very important to gain structural insights of all the phases. But, to get a single crystal structure after each dynamic process is quite difficult. Thus, using single crystal X-ray diffraction, PXRD, other modern structure determining tools and theoretical methods,⁷⁷ neutral N-donor ligand based MOFs might show new functions because of their structural flexibility, for the development of new functional materials for various potential applications.

Acknowledgements

B. M. is thankful to CSIR for research fellowship, while IISER Pune is acknowledged for the same by B. M., A. V. D.; DST (Project no. GAP/DST/CHE-12-0083) and DST-FIST (SR/FST/CSII-023/2012) are acknowledged for generous financial support.

Notes and references

- H.-C. Zhou, J. R. Long and O. M. Yaghi, *Chem. Rev.*, 2012, **112**, 673.
- J.-R. Li, J. Sculley and H.-C. Zhou, *Chem. Rev.*, 2012, **112**, 869.
- B. Chen, Y. Yang, F. Zapata, G. Lin, G. Qian and E. B. Lobkovsky, *Adv. Mater.*, 2007, **19**, 1693.
- B. Chen, L. Wang, F. Zapata, G. Qian and E. B. Lobkovsky, *J. Am. Chem. Soc.*, 2008, **130**, 6718.
- P. Ramaswamy, N. E. Wong and G. K. H. Shimizu, *Chem. Soc. Rev.*, 2014, **43**, 5913.
- D. N. Dybtsev, H. Chun and K. Kim, *Angew. Chem., Int. Ed.*, 2004, **43**, 5033.
- S. Horike, S. Shimomura and S. Kitagawa, *Nat. Chem.*, 2009, **1**, 695.
- C. Serre, C. Mellot-Draznieks, S. Surble, N. Audebrand, Y. Filinchuk and G. Férey, *Science*, 2007, **315**, 1828.
- G. Kumar Kole and J. J. Vittal, *Chem. Soc. Rev.*, 2013, **42**, 1755.
- A. Schneemann, V. Bon, I. Schwedler, I. Senkovska, S. Kaskel and R. A. Fischer, *Chem. Soc. Rev.*, 2014, **43**, 6062.
- Z. Chang, D.-H. Yang, J. Xu, T.-L. Hu and X.-H. Bu, *Adv. Mater.*, 2015, **27**, 5432–5441.
- M. C. Das and P. K. Bharadwaj, *J. Am. Chem. Soc.*, 2009, **131**, 10942.
- Y.-Q. Chen, G.-R. Li, Z. Chang, Y.-K. Qu, Y.-H. Zhang and X.-H. Bu, *Chem. Sci.*, 2013, **4**, 3678.
- J.-P. Ma, Y. Yu and Y.-B. Dong, *Chem. Commun.*, 2012, **48**, 2946.
- H. Fei, M. R. Bresler and S. R. J. Oliver, *J. Am. Chem. Soc.*, 2011, **133**, 11110.
- H. Fei, C. S. Han, J. C. Robins and S. R. J. Oliver, *Chem. Mater.*, 2013, **25**, 647.
- X. Li, H. Xu, F. Kong and R. Wang, *Angew. Chem., Int. Ed.*, 2013, **52**, 13769.
- B. Manna, B. Joarder, A. V. Desai, A. Karmakar and S. K. Ghosh, *Chem. – Eur. J.*, 2014, **20**, 12399.
- K. Biradha, K. V. Domasevitch, B. Moulton, C. Seward and M. J. Zaworotko, *Chem. Commun.*, 1999, 1327.
- M. C. Das, S. K. Ghosh, E. C. Sañudo and P. K. Bharadwaj, *Dalton Trans.*, 2009, 1644.
- W. L. Leong and J. J. Vittal, *Chem. Rev.*, 2011, **111**, 688.
- A. Karmakar, A. V. Desai and S. K. Ghosh, *Coord. Chem. Rev.*, 2015, DOI: 10.1016/j.ccr.2015.08.007.
- C.-P. Li, J. Chen, C.-S. Liub and M. Du, *Chem. Commun.*, 2015, **51**, 2768.
- J.-P. Zhang, P.-Q. Liao, H.-L. Zhou, R.-B. Lin and X.-M. Chen, *Chem. Soc. Rev.*, 2014, **43**, 5789.
- O.-S. Jung, Y. J. Kim, Y.-A. Lee, H. K. Chae, H. G. Jang and J. Hong, *Inorg. Chem.*, 2001, **40**, 2105.
- S. Muthu, J. H. K. Yip and J. J. Vittal, *J. Chem. Soc., Dalton Trans.*, 2002, 4561.
- S. Kitagawa, R. Kitaura and S.-i. Noro, *Angew. Chem., Int. Ed.*, 2004, **43**, 2334.
- O. M. Yaghi and H. Li, *J. Am. Chem. Soc.*, 1996, **118**, 295.
- K. S. Min and M. P. Suh, *J. Am. Chem. Soc.*, 2000, **122**, 6834.
- O.-S. Jung, Y. J. Kim, Y.-A. Lee, J. K. Park and H. K. Chae, *J. Am. Chem. Soc.*, 2000, **122**, 9921.
- S.-I. Noro, R. Kitaura, M. Kondo, S. Kitagawa, T. Ishii, H. Matsuzaka and M. Yamashita, *J. Am. Chem. Soc.*, 2002, **124**, 2568.
- T. K. Maji, R. Matsuda and S. Kitagawa, *Nat. Mater.*, 2007, **6**, 142.
- B. Manna, A. K. Chaudhari, B. Joarder, A. Karmakar and S. K. Ghosh, *Angew. Chem., Int. Ed.*, 2013, **52**, 998.
- C. P. Li, J. Guo and M. Du, *Inorg. Chem. Commun.*, 2013, **38**, 70.
- S. Hou, Q.-K. Liu, J.-P. Ma and Y.-B. Dong, *Inorg. Chem.*, 2013, **52**, 3225.
- B. Manna, S. Singh, A. Karmakar, A. V. Desai and S. K. Ghosh, *Inorg. Chem.*, 2015, **54**, 110.
- A. Karmakar, A. V. Desai, B. Manna, B. Joarder and S. K. Ghosh, *Chem. – Eur. J.*, 2015, **21**, 7071.



- 38 J. J. Vittal, *Coord. Chem. Rev.*, 2007, **251**, 1781.
- 39 L. Carlucci, G. Ciani, M. Moret, D. M. Proserpio and S. Rizzato, *Angew. Chem., Int. Ed.*, 2000, **39**, 1506.
- 40 K. Biradha, Y. Hongo and M. Fujita, *Angew. Chem., Int. Ed.*, 2002, **41**, 3395.
- 41 K. Biradha and M. Fujita, *Angew. Chem., Int. Ed.*, 2002, **41**, 3392.
- 42 K. Uemura, S. Kitagawa, K. Fukui and K. Saito, *J. Am. Chem. Soc.*, 2004, **126**, 3817.
- 43 Y. Inokuma, T. Arai and M. Fujita, *Nat. Chem.*, 2010, **2**, 780.
- 44 W. M. Bloch and C. J. Sumby, *Chem. Commun.*, 2012, **48**, 2534.
- 45 W. M. Bloch and C. J. Sumby, *Eur. J. Inorg. Chem.*, 2015, **2015**, 3723.
- 46 A. Karmakar, B. Manna, A. V. Desai, B. Joarder and S. K. Ghosh, *Inorg. Chem.*, 2014, **53**, 12225.
- 47 B. Manna, A. V. Desai, N. Kumar, A. Karmakar and S. K. Ghosh, *CrystEngComm*, 2015, DOI: 10.1039/C5CE00139K.
- 48 A. Malik, P. Peedikakkal and J. J. Vittal, *Cryst. Growth Des.*, 2011, **11**, 4697.
- 49 M. H. Zeng, S. Hu, Q. Chen, G. Xie, Q. Shuai, S. L. Gao and L. Y. Tang, *Inorg. Chem.*, 2009, **48**, 7070.
- 50 D. Bradshaw, J. E. Warren and M. J. Rosseinsky, *Science*, 2007, **315**, 977.
- 51 A. Aslani and A. Morsali, *Chem. Commun.*, 2008, 3402.
- 52 C.-F. Zhuang, J. Zhang, Q. Wang, Z.-H. Chu, D. Fenske and C.-Y. Su, *Chem. – Eur. J.*, 2009, **15**, 7578.
- 53 X.-M. Liu, B.-Y. Wang, W. Xue, L.-H. Xie, W.-X. Zhang, X.-N. Cheng and X.-M. Chen, *Dalton Trans.*, 2012, **41**, 13741.
- 54 J. L. C. Rowsell, E. C. Spencer, J. Eckert, J. A. K. Howard and O. M. Yaghi, *Science*, 2005, **309**, 1350.
- 55 H. Sato, W. Kosaka, R. Matsuda, A. Hori, Y. Hijikata, R. V. Belosludov, S. Sakaki, M. Takata and S. Kitagawa, *Science*, 2014, **343**, 167.
- 56 D. Li and K. Kaneko, *Chem. Phys. Lett.*, 2001, **334**, 50.
- 57 A. Kondo, H. Noguchi, S. Ohnishi, H. Kajiro, A. Tohdoh, Y. Hattori, W.-C. Xu, H. Tanaka, H. Kanoh and K. Kaneko, *Nano Lett.*, 2006, **6**, 2581.
- 58 A. Kondo, H. Noguchi, L. Carlucci, D. M. Proserpio, G. Ciani, H. Kajiro, T. Ohba, H. Kanoh and K. Kaneko, *J. Am. Chem. Soc.*, 2007, **129**, 12362.
- 59 A. Kondo, H. Kajiro, H. Noguchi, L. Carlucci, D. M. Proserpio, G. Ciani, K. Kato, M. Takata, H. Seki, M. Sakamoto, Y. Hattori, F. Okino, K. Maeda, T. Ohba, K. Kaneko and H. Kanoh, *J. Am. Chem. Soc.*, 2011, **133**, 10512.
- 60 R. Kotani, A. Kondo and K. Maeda, *Chem. Commun.*, 2012, **48**, 11316.
- 61 I.-H. Park, A. Chanthapally, Z. Zhang, S. S. Lee, M. J. Zaworotko and J. J. Vittal, *Angew. Chem., Int. Ed.*, 2014, **53**, 414.
- 62 M. Nagarathinam and J. J. Vittal, *Angew. Chem., Int. Ed.*, 2006, **45**, 4337.
- 63 M. Nagarathinam and J. J. Vittal, *Chem. Commun.*, 2008, 438.
- 64 A. M. P. Peedikakkal and J. J. Vittal, *Inorg. Chem.*, 2010, **49**, 10.
- 65 R. Medishetty, T. T. S. Yap, L. L. Koh and J. J. Vittal, *Chem. Commun.*, 2013, **49**, 9567.
- 66 R. Medishetty, S. C. Sahoo, C. E. Mulijanto, P. Naumov and J. J. Vittal, *Chem. Mater.*, 2015, **27**, 1821.
- 67 G. Ferrey and C. Serre, *Chem. Soc. Rev.*, 2009, **38**, 1380.
- 68 S. Tominaka, H. Hamoudi, T. Suga, T. D. Bennett, A. B. Cairns and A. K. Cheetham, *Chem. Sci.*, 2015, **6**, 1465.
- 69 G. J. Halder, C. J. Kepert, B. Moubaraki, K. S. Murray and J. D. Cashion, *Science*, 2002, **298**, 1762.
- 70 S. M. Neville, B. Moubaraki, K. S. Murray and C. J. Kepert, *Angew. Chem., Int. Ed.*, 2007, **46**, 2059.
- 71 S. M. Neville, G. J. Halder, K. W. Chapman, M. B. Duriska, P. D. Southon, J. D. Cashion, J.-F. Letard, B. Moubaraki, K. S. Murray and C. J. Kepert, *J. Am. Chem. Soc.*, 2008, **130**, 2869.
- 72 G. J. Halder, K. W. Chapman, S. M. Neville, B. Moubaraki, K. S. Murray, J.-F. Letard and C. J. Kepert, *J. Am. Chem. Soc.*, 2008, **130**, 17552.
- 73 S. M. Neville, G. J. Halder, K. W. Chapman, M. B. Duriska, B. Moubaraki, K. S. Murray and C. J. Kepert, *J. Am. Chem. Soc.*, 2009, **131**, 12106.
- 74 S.-I. Noro, Y. Hijikata, M. Inukai, T. Fukushima, S. Horike, M. Higuchi, S. Kitagawa, T. Akutagawa and T. Nakamura, *Inorg. Chem.*, 2013, **52**, 280.
- 75 P. Nugent, Y. Belmabkhout, S. D. Burd, A. J. Cairns, R. Luebke, K. Forrest, T. Pham, S. Ma, B. Space, L. Wojtas, M. Eddaoudi and M. J. Zaworotko, *Nature*, 2013, **495**, 80.
- 76 J.-H. Wang, M. Li and D. Li, *Chem. Sci.*, 2013, **4**, 1793.
- 77 A. U. Ortiz, A. Boutin, A. H. Fuchs and F.-X. Coudert, *Phys. Rev. Lett.*, 2012, **109**, 195502.

