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In situ spectroscopic investigations of MoO_x/Fe₂O₃ catalysts for the selective oxidation of methanol†

Catherine Brookes,^{*ab} Michael Bowker,^{ab} Emma K. Gibson,^{ac} Diego Gianolio,^d
Khaled M. H. Mohammed,^{ace} Stephen Parry,^d Scott M. Rogers,^{ac} Ian P. Silverwood^{ac}
and Peter P. Wells^{*ac}

Multicomponent oxide shell@core catalysts have been prepared, affording overlayers of MoO_x on Fe₂O₃. This design approach allows bulk characterization techniques, such as X-ray Absorption Fine Structure (XAFS), to provide surface sensitive information. Coupling this approach with *in situ* methodologies provides insights during crucial catalytic processes. Calcination studies were followed by a combination of XAFS and Raman, and demonstrate that amorphous multi-layers of MoO_x are first converted to MoO₃ before formation of Fe₂(MoO₄)₃. However, a single overlayer of O_h Mo units remains at the surface at all times. *In situ* catalysis studies during formaldehyde production identified that Mo⁶⁺ was present throughout, confirming that gas phase oxygen transfer to molybdenum is rapid under reaction conditions. Reduction studies in the presence of MeOH resulted in the formation of reduced Mo–Mo clusters with a bonding distance of 2.6 Å. It is proposed that the presence of the clusters indicates that the selective conversion of MeOH to formaldehyde requires multiple Mo sites.

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Introduction

Multi-component oxides have a wide range of catalytic applications, including selective alcohol oxidation and oxidative dehydrogenation of small chain alkanes.^{1–4} The activity of these heterogeneous catalysis materials is determined by the structural properties of their uppermost layer, which often vary considerably from the bulk structure.^{5,6} It has recently been demonstrated that the surface layer of Fe₂(MoO₄)₃ is significantly different to the bulk structure; the surface layer is composed of O_h Mo units, compared to the T_d Mo centres found in Fe₂(MoO₄)₃.^{6,7} This is a key finding, as Fe₂(MoO₄)₃ is a catalyst of significant industrial importance, being widely used in the selective oxidation of methanol to formaldehyde.^{5,8–12} Formaldehyde synthesis is a crucial industrial process, producing in excess of 6 million tonnes

annually, and widely used as a building block to higher value products, *e.g.* resins, and adhesives.

Unravelling the nature of the active site in commercial Fe₂(MoO₄)₃ catalysts has always proved challenging, as excess Mo is incorporated into the synthesis procedure.^{13–15} Any Mo lost through sublimation has the potential to leave a vacant, unselective Fe site,¹⁶ and so the excess Mo is present to counter this eventuality.¹⁷ However, this approach is shown to leave a system containing multiple distinct phases including both MoO₃ and Fe₂(MoO₄)₃, with MoO₃ shown to be present as particles that segregate to the surface.¹⁸ With a multitude of phases, this makes it difficult to assess the nature of the active site, with any *in situ* spectroscopic study probing a variety of surfaces and many spectator species. Conventional *in situ* X-ray diffraction (XRD) or X-ray absorption fine structure (XAFS) studies of bulk Fe₂(MoO₄)₃ catalysts are unable to identify how the surface layer changes during important catalytic processes,^{19,20} as the surface layer makes up only a small proportion of the overall Mo sites. Our recent study⁷ on single MoO_x overlayers on Fe₂O₃ demonstrate a comparable performance to Fe₂(MoO₄)₃ catalysts. XAFS studies are element specific and provide a per atom average of all environments. However, when Mo is exclusively at the surface, XAFS is able to provide surface-sensitive information. Subsequently, when XAFS is applied under relevant *in situ/operando* conditions, it can provide accurate information about the nature of the active site in MoO_x/Fe₂O₃ catalysts. The work herein demonstrates how model MoO_x/Fe₂O₃ catalysts can be studied *in situ*

^a UK Catalysis Hub, Research Complex at Harwell, Rutherford Appleton Laboratory, Harwell, Oxon, OX11 0FA, UK. E-mail: catherine.brookes@rc-harwell.ac.uk, peter.wells@rc-harwell.ac.uk

^b Cardiff Catalysis Institute, School of Chemistry, Cardiff University, Park Place, Cardiff, CF10 3AT, UK

^c Department of Chemistry, University College London, 20 Gordon St., London, WC1H 0AJ, UK

^d Diamond Light Source, Harwell Science and Innovation Campus, Didcot, Oxon, OX11 0DE, UK

^e Chemistry Department, Faculty of Science, Sohag University, Sohag, P.O.B 82524, Egypt

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for their preparation, reaction and deactivation, in order to shed light on how the same processes occur in conventional iron molybdate catalysts. More specifically, we address how the reaction mechanism proceeds at the surface, and how the catalyst is able to regenerate for further reaction and maintained selectivity.

Experimental

Catalyst preparation

$\text{MoO}_x/\text{Fe}_2\text{O}_3$ was prepared using the incipient wetness technique. Commercial Fe_2O_3 (Sigma Aldrich, <50 nm) was first calcined at 500 °C for 3 hours, before adding the desired amount of aqueous ammonium heptamolybdate for 1, 3 or 6 monolayers (MLS) $\text{MoO}_x/\text{Fe}_2\text{O}_3$. The samples were dried at 120 °C for 24 hours, before being calcined at 500 °C for 24 hours in a muffle furnace. For subsequent reduction studies, the calcined materials were reduced in a tube furnace under a methanol/He flow at 350 °C for three hours. Bulk phase samples were synthesised to be used as reference standards in the analysis. Stoichiometric $\text{Fe}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ (Fe(III), Sigma Aldrich >98%) was added drop wise to a stirring solution of $(\text{NH}_4)_6\text{Mo}_7\text{O}_{24} \cdot 4\text{H}_2\text{O}$ (Fluka Analytical, >99%) previously acidified to pH 2 with dilute HNO_3 . The resulting mixture was heated to 90 °C for ~1 hour until a yellow sludge remained. This was left to air dry overnight, after which it was dried at 120 °C for 24 hours, followed by calcination at 500 °C for 48 hours.

Physical characterization

BET measurements were taken for the commercial Fe_2O_3 and prepared catalysts using a Micromeritics Gemini surface area analyzer. Raman measurements were carried out using a Renishaw Raman microscope with an 830 nm laser, over a wavenumber range of 100–1200 cm^{-1} . Typical measurements used a 0.1% laser power, with 4 accumulations at 10 s exposure time for each. XRD of reference standards were performed on a Panalytical X'pert pro analyser with Cu K α radiation. The IR data were recorded using a Thermo Scientific Nicolet iS10 FT-IR Spectrometer using an ATR accessory in the range of 500–2500 cm^{-1} . Catalyst morphology and homogeneity were examined by both TEM (JEOL JEM 2100 EM) and SEM (JEOL JSM- 6610LV).

Catalytic testing

Reactor data were obtained by use of a CATLAB reactor (Hiden Ltd, Warrington UK). For the catalytic reaction, 1 μl of liquid methanol was injected every 2 minutes into a flow of 10% O_2/He , at a flow rate of 30 ml min^{-1} . The products were determined by the online mass spectrometer integrated into the CATLAB system. For the TPD, approximately 6 injections of 1 μl methanol were dosed onto the catalyst, at ambient temperature in a flow of 30 ml min^{-1} of He. This was followed by ramping the temperature to 400 °C at a rate of

8 °C min^{-1} , whilst simultaneously monitoring the products formed by mass spectrometry.

XAFS measurements

Mo k-edge XAFS studies were carried out on the B18 beamline at the Diamond Light Source, Didcot, UK. Measurements were performed using a QEXAFS set-up with a fast-scanning Si(111) double crystal monochromator. The time resolution of the spectra reported herein was 5 min per spectrum ($k_{\text{max}} = 14$), on average three scans were acquired to improve the signal to noise level of the data. *Ex situ* measurements were performed for the samples of $\text{Fe}_2(\text{MoO}_4)_3$, MoO_3 and dried 3 and 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$, to be used as reference standards for linear combination analysis in XANES interpretations. All *ex situ* samples were diluted with cellulose and pressed into pellets to optimise the effective edge-step of the XAFS data and measured in transmission mode using ion chamber detectors. All transmission XAFS spectra were acquired concurrently with a Mo foil placed between I_t and I_{ref} . XAS data processing and EXAFS analysis were performed using IFEFFIT²¹ with the Horae package²² (Athena and Artemis). The amplitude reduction factor, s_0^2 , was derived from EXAFS data analysis of a known Mo reference compound, MoO_3 , (with known coordination numbers which were fixed during analysis) to be 0.82, which was used as a fixed input parameter.

In situ calcination and reduction studies were performed in transmission mode with a quartz capillary (OD 6 mm, wall thickness 250 μm) microreactor (Fig. 1). The microreactor is equipped with a gas supply system, integrated heating system and a Cirrus 100 quadrupole mass spectrometer to monitor the outlet gas composition online under various conditions. For the calcination studies, approximately 0.5 g of dried $\text{MoO}_x/\text{Fe}_2\text{O}_3$ was heated using a staircase heating procedure with a ramp rate of 12 °C min^{-1} between 20 °C increments. The temperature was held for ten minutes at each increment point to 650 °C. XAFS data were recorded throughout the experiment, at a typical rate of 1 min per scan. On reaching the maximum temperature, the sample was held isothermally for 30 minutes to ensure no further changes occurred in the



Fig. 1 *In situ* quartz capillary microreactor for XAFS studies. (A) Hot air blower, (B) I_0 , (C) I_t , and (D) fluorescence detector.



structure, after which it was cooled to room temperature. In addition to the *in situ* calcination described, the catalyst was ramped in temperature and held isothermally for 120 minutes at 250 °C. XAFS was monitored as MeOH in O₂/He was passed over the catalyst bed, inducing a reaction at the surface.

Reduction studies were carried out in a MeOH/He flow whilst heating to 350 °C at a rate of 12 °C min⁻¹. XAFS spectra were collected throughout (~1 min per scan). Upon reaching 350 °C, the gas flow mixture was switched to pure He, and the sample cooled to room temperature.

Combined *in situ* XAFS/Raman studies

XAFS/Raman studies were performed using the same micro-reactor described, coupled with a custom-built optic-fibre-coupled, confocal Raman probe system (Fig. 2). The system uses a Renishaw RA100 spectrometer, with a Peltier cooled CCD detector and 1200 line grating, and a 300 mW, 785 nm diode laser (Renishaw HP785NIR). Laser light from the diode source was directed into the Raman probe using a multimode fibre optic, where lenses expanded and collimated the light. A ×20 magnification lens at the exit aperture of the probe focused the laser light to a ~500 μm spot at the sample. Edge filters (cut-off 100 nm above the laser line wavelength) were used to reject the laser light in the back reflected signal into the probe, and a condensing lens was used to focus this signal into a second multimode fibre optic which directed the signal into the RA100 spectrometer. Slits were used to attenuate the laser power on the sample to ~15 mW, to prevent sample damage and saturation of the detector. The spectrometer was calibrated using the emission lines from a Ne lamp, and checked using the main Raman band from a silicon wafer, setting the centre of this peak to 520 cm⁻¹. Spectra were collected for 20 × 1 second scans using 'static' mode, with the grating angle fixed to illuminate the Raman spectrum across the CCD. The techniques were combined to follow the calcination study (as described previously) with the XAFS geometry adjusted to fluorescence mode. The change in acquisition mode to fluorescence allowed for undiluted materials to be used, in order to maximise Raman signal from the sample.

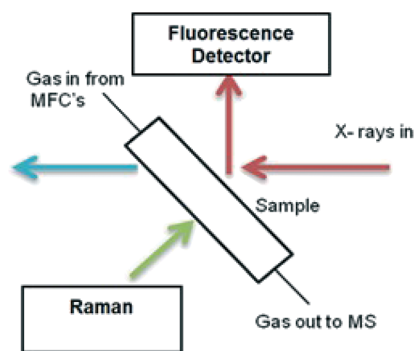


Fig. 2 Schematic representation of the *in situ* XAFS/Raman set up performed in fluorescence mode.

Results and discussion

In situ calcination studies

Prior to any experiments performed on the beamline, all catalysts were characterized by a variety of techniques. SEM-EDX mapping was performed to confirm the Mo loading, and to show that the molybdenum was well-dispersed across the entire catalyst (Fig. S1†). TEM analysis established the Mo to be present at the surface through EDX line scans (Fig. S2†). With knowledge of catalyst composition and homogeneity, vibrational studies were then able to provide insight into the oxide phases present with varying calcination temperature *ex situ*. Fig. S3–4† detail Raman spectroscopic studies for the 3 and 6 ML MoO_x/Fe₂O₃ catalysts. Both show an obvious change from a MoO₃ phase formed at elevated temperatures (>300 °C) to Fe₂(MoO₄)₃ by 500 °C.

The Mo K-edge XAFS data of the 3 and 6 ML MoO_x/Fe₂O₃ samples calcined to 120, 300, 400, 500 and 600 °C have been previously measured *ex situ* and reported elsewhere.⁷ Although *ex situ* analysis provides a good indication of the catalyst composition at various pre-defined temperatures, it is limited in that it is unable to detect any transient phases formed prior to those identified at the selected set points. *In situ* studies enable the observation of successive changes within the catalyst structure under variable experimental conditions. The interpretation of features in the Mo XANES spectra involves the assignment of the pre-edge peak at ~19995 eV and the absorption edge at 20 010 eV. The pre-edge peak is for the dipole forbidden/quadrupole allowed 1s–4d transition,²³ and is most intense for tetrahedral Mo. However, it is also evident, but less intense, for structures with distorted octahedral geometry.^{24,25} The absorption edge at 20 010 eV is assigned to the dipole allowed 1s–5p transition, and is a characteristic feature of Mo species with octahedral/distorted octahedral geometry. Results for the *in situ* calcination study are presented in Fig. 3–5. Data for the 3 and 6 ML MoO_x/Fe₂O₃ catalysts are conclusive, with both showing a substantial enhancement of the Mo pre-edge with temperature increase. This is indicative of a transition from O_h to T_d geometry. The effect was most pronounced for the 6 ML MoO_x/Fe₂O₃ catalyst, as expected for the higher weight loading of Mo, which forms the greatest fraction of the Fe₂(MoO₄)₃ T_d phase. The experiments involved steadily increasing the temperature until no further changes were observed in the XANES. LCA were performed to assess the amount of MoO_x, MoO₃ and Fe₂(MoO₄)₃ present within the sample upon varying the temperature, with results presented in Fig. 4–5.

For the calcination study of 3 ML MoO_x/Fe₂O₃ (Fig. 4), upon ramping from 300 °C, there is an immediate and steady decrease in the MoO_x phase initially dosed at the surface. Parallel to this, there is also a slow but steady rise in MoO₃ formation. At this point, the amount of Fe₂(MoO₄)₃ within the fit does not rise significantly. By 520 °C, a discernible loss of MoO₃ occurs, dropping by 6%. This point marks the onset of Fe₂(MoO₄)₃ formation. The results imply that formation of MoO₃ and the reaction with the Fe₂O₃ core is responsible for





Fig. 3 Normalised XANES data for 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ *in situ* calcination study. Approximately 0.5 g of the dried material for 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ was mixed with BN and loaded into the sample capillary. A 10% O_2/He mixture was passed over the catalyst whilst heating to 640 °C at a ramp rate of 12 °C min^{-1} . Temperatures from 300 °C to 640 °C are shown at 20 °C intervals. No changes were seen in the XANES below 300 °C.

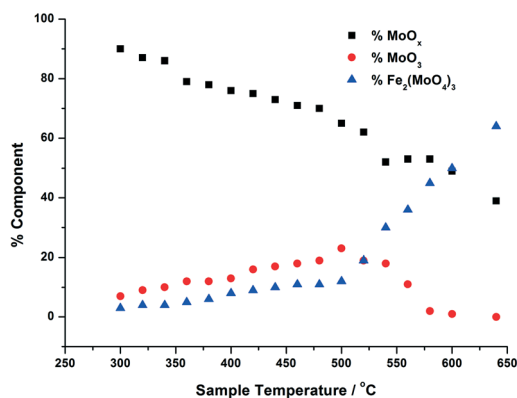


Fig. 4 Phase composition of MoO_x , MoO_3 , $\text{Fe}_2(\text{MoO}_4)_3$ with increasing calcination temperature, for 3 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$. The calcination was performed as in Fig. 3, and the phase compositions were determined using linear combination analysis of MoO_3 and $\text{Fe}_2(\text{MoO}_4)_3$ (Sigma Aldrich) and the dried 3 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$.

the formation of $T_d \text{Fe}_2(\text{MoO}_4)_3$. At the highest temperature of 650 °C there is no evidence of MoO_3 , and $\text{Fe}_2(\text{MoO}_4)_3$ dominates with 64% of the total Mo dosed. It is important to note, that even at this temperature LCA demonstrates a 1 ML equivalent of surface MoO_x still present, which fails to be incorporated into the $\text{Fe}_2(\text{MoO}_4)_3$ phase. The stability of this overlayer has previously been established,^{7,26} as it still remains after calcination at 600 °C for 24 hours. A complementary study was performed for the 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalyst. It is important to highlight that a relative comparison could not be made to the 1 ML case since here we see no change in phase across the calcination range (Fig. S5†). The trends observed for the 6 ML catalyst (Fig. 5) are similar in nature to those discussed for 3 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$, however, there are subtle variations. Again, the MoO_3 is seen to primarily form, but in this case develops more rapidly at lower temperatures, with significant formation at 300 °C. For the 6 ML coverage there is double the amount of Mo dosed,



Fig. 5 Phase composition of MoO_x , MoO_3 , $\text{Fe}_2(\text{MoO}_4)_3$ with increasing calcination temperature, for 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$. The calcination was performed as in Fig. 3, and the phase compositions were determined using linear combination analysis of MoO_3 and $\text{Fe}_2(\text{MoO}_4)_3$ (Sigma Aldrich) and the dried 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$.

therefore the proportion of MoO_3 formed is higher than that in the 3 ML catalyst. By 360 °C, XANES analysis demonstrates 20% MoO_3 , with only 8% $\text{Fe}_2(\text{MoO}_4)_3$ formed. The amount of $\text{Fe}_2(\text{MoO}_4)_3$ is nearly constant between 300 and 500 °C, and only begins to sharply increase after 500 °C, and coincides with the decrease in MoO_3 as it reacts with bulk Fe_2O_3 . As was noted for the 3 ML sample, at the end of the calcination, we are left with a resultant 1 ML equivalent of MoO_x , remaining segregated at the surface of the Fe_2O_3 .

Fig. 6 presents the Raman data obtained from the *in situ* XAFS/Raman set up. Results reiterate the changes seen through XANES for the 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$. For the purpose of this work, the 6 ML loading is discussed due to its higher Raman signal, however, results were also consistent for the 3 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalyst. MoO_3 can be observed through the presence of bands at 820 and 990 cm^{-1} , assigned to the Mo–O–Mo bridge and Mo=O terminal stretch, respectively.⁵ On increasing the temperature these bands diminish above

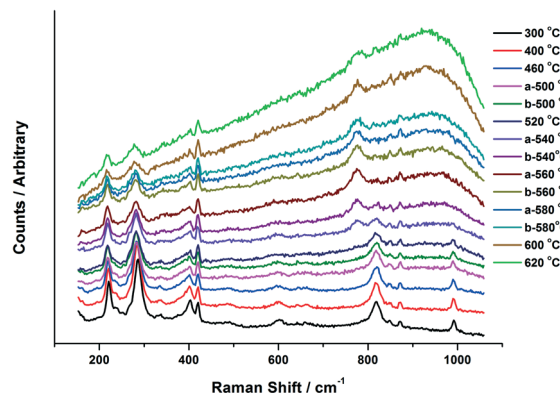


Fig. 6 Raman data collected *in situ* with combined XAFS/Raman for 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$. The catalyst was ramped to 640 °C at a rate of 12 °C min^{-1} . Spectra were collected for 20 × 1 second scans using 'static' mode, with the grating angle fixed to illuminate the Raman spectrum across the CCD. Temperatures from 300 °C to 620 °C are shown, those labelled (a) are the initial heating ramp, those labelled (b) represent the cooling cycle.



500 °C, which is in agreement with the results seen in the XANES analysis. By 540 °C, a new band appears at 780 cm⁻¹, indicative of asymmetric Mo=O stretching in isolated MoO₄ sites in Fe₂(MoO₄)₃.⁵ The intensity of this band increases with further rises in temperature, with the bands at 820 and 990 cm⁻¹ (corresponding to MoO₃) decreasing. This is as a result of the conversion of MoO₃ to Fe₂(MoO₄)₃, through reaction with the Fe₂O₃ support. Above 600 °C, the peak intensity readily diminishes. These effects can be ascribed to the increased mobility of cationic species at higher temperature. On cooling, the bands associated with Fe₂(MoO₄)₃ were visible once again as order within the sample is regained.

The combined XAFS/Raman experiment was able to provide the following information, which the *ex situ* experiments alone were unable to provide:

i) The production of MoO₃ progressively forms with rising calcination temperature. This was seen in both XANES and Raman spectra for the catalysts studied. Only once the levels of MoO₃ plateau, does Fe₂(MoO₄)₃ production commence. This would infer that MoO₃ is a prerequisite to Fe₂(MoO₄)₃ formation.

ii) The onset of substantial MoO₃ formation occurs significantly earlier in 6 ML MoO_x/Fe₂O₃ than 3 ML MoO_x/Fe₂O₃, however both catalysts could not form any Fe₂(MoO₄)₃ until 500 °C. It is believed therefore, that this is the thermal input required to form the mixed oxide.

The Raman studies were supported by complimentary *in situ* DRIFTS measurements of the 6 ML MoO_x/Fe₂O₃ during calcination. Results can be referred to in Fig. S6†. With these studies, we have been able to demonstrate the use of combined *in situ* techniques to observe changes in catalyst structure under preparation conditions.

In situ catalysis studies

The nature of Mo during methanol oxidation was assessed with an *in situ* XAFS investigation. Fig. 7 shows the data obtained throughout a 120 minute time period, for the 6 ML MoO_x/Fe₂O₃ catalyst. No obvious changes were observed in the absorption edge position, whilst T_d co-ordination was maintained during the course of the reaction. The XAFS investigations found that the Mo environment was preserved under reaction conditions, enabling the sustained production of formaldehyde, as shown through the mass spectrometer data collected online for the duration of the experiment (Fig. 8). This is a direct consequence of the catalysts ability in maintaining its Mo(vi) state,^{10,27,28} through re-oxidation from gas phase oxygen. Mo(iv) may exist momentarily, but is unable to be seen due to the time scale of these studies.

In situ reduction-oxidation studies

Although highly regarded for its efficiency in the reaction of methanol to formaldehyde, there is a somewhat limited knowledge regarding the reaction mechanisms involved at the catalyst surface of iron molybdate. It is agreed that during the reaction, partial reduction of the surface occurs,

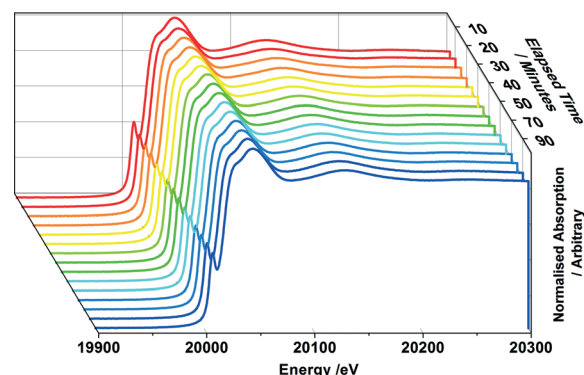
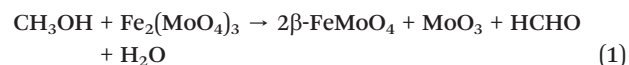


Fig. 7 XANES data over a 90 minute period for 6 ML MoO_x/Fe₂O₃ when reacted isothermally at 250 °C under a continuous flow of MeOH/O₂. XAFS data was collected continuously using 1 minute scans, whilst Mass Spectrometer data was recorded simultaneously, to ensure a reaction was occurring.

which is then counteracted by a quick regeneration through gas phase oxygen.^{29,30} It could be that Mo(iv) exists temporarily, however, it has not been possible to isolate this species under normal reaction conditions. A reduction study would be able to provide further information about the redox cycling of the Mo species.

On bulk Fe₂(MoO₄)₃,²⁹ the reduction mechanism has been reported to occur as follows, resulting in a mixed phase catalyst:



Which then re-oxidises as follows;



Zhang *et al.*³¹ used TPR to study the reduction process on Fe₂(MoO₄)₃ with excess MoO₃. They define peaks ascribed to the transformation of Fe₂(MoO₄)₃ to β-FeMoO₄ and Mo₄O₁₁, MoO₃ to MoO₂ and β-FeMoO₄ to Fe₂Mo₃O₈ and Fe₃O₄. Beale



Fig. 8 Online Mass Spectrometry data for 6 ML MoO_x/Fe₂O₃, when reacted isothermally in continuous flow under MeOH/O₂ at 250 °C (see Fig. 7 for details). Selectivity to formaldehyde was 90%. The data has not been normalised to the fragmentation pattern of products produced.



et al. enforce this idea, using combined *in situ* techniques,²⁹ specifically WAXS, XANES and UV-Vis, to investigate $\text{Fe}_2(\text{MoO}_4)_3$ in more detail. Under redox conditions, they were able to confirm the reduction to $\beta\text{-FeMoO}_4$ and MoO_3 , which can further reduce to MoO_2 .

With a multitude of different phases forming under reduction, characterisation becomes complicated. Further to this, bulk techniques such as Raman spectroscopy are limited due a number of reasons;

i) They are unable to explicitly characterise the surface layer.
 ii) Problems arise in studying many of the reduced phase catalysts by spectroscopy, as a result of their dark colour and absorbing characteristics.

iii) Previous characterisation performed *ex situ* has revealed that catalysts re-oxidise when left under atmospheric conditions. UV-Vis studies (Fig. S7†) of the post reacted and reduced samples show a mixed $\text{Mo}^{5+/6+}$ oxidation state, an indication that the catalyst has partially re-oxidised back to its original state. This is further accelerated under the heat of the focussed Raman laser.

In this section we investigate the mechanisms occurring at the surface of FeMo based catalysts under reaction with MeOH/He . Initial *ex situ* studies probed the reduction process in bulk $\text{Fe}_2(\text{MoO}_4)_3$. Data presented through XAFS analysis (Fig. S10†) clearly evidenced the formation of MoO_2 and FeMoO_4 . This was also ascertained through a variety of techniques, including Raman and XRD. (Fig. S8–9†).

In studying $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalysts, this allows us to exclusively probe the active topmost layer, without the added complication of secondary phases such as MoO_2 and $\beta\text{-FeMoO}_4$, which we have shown arise through the reduction of bulk $\text{Fe}_2(\text{MoO}_4)_3$. 1, 3 and 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalysts have been investigated under reduction with MeOH/He . Fig. 9 details the extent of reduction as a function of temperature for each of these monolayer coverages, acquired through XANES analysis. Since it was not possible to source appropriate references to perform satisfactory LCA, results were ascertained through observing the shift in edge position between Mo(VI) from the original post reduced catalyst, to Mo(IV) referenced from MoO_2 . Based on these results, we demonstrate an edge shift of approximately 70% towards Mo(IV) for all coverages. Further characterization through Raman and XRD analysis, was unable to establish isolated MoO_2 , as is the case for bulk $\text{Fe}_2(\text{MoO}_4)_3$ (Fig. S8–9†).

To elucidate this reduction process further, EXAFS analysis has played a key role. The end state k^2 -weighted Fourier transform data for the reduced samples of the 1 and 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalysts are shown in Fig. 10. Results for the 3 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalyst were in accordance with those seen for 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ (Fig. S11†), so comparisons herein will be made between the 1 and 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalysts for clarity. The obtained fits for all catalysts are comprised of 5 scattering paths; 2 Mo–O, Mo–Mo and 2 Mo–M, where $\text{M}=\text{Fe}$ or Mo . Two oxygen scattering paths feature at 1.78 and 2.02 Å, indicative of reduced forms of Mo, as well as the distorted nature of the Mo octahedra. In addition to this, all

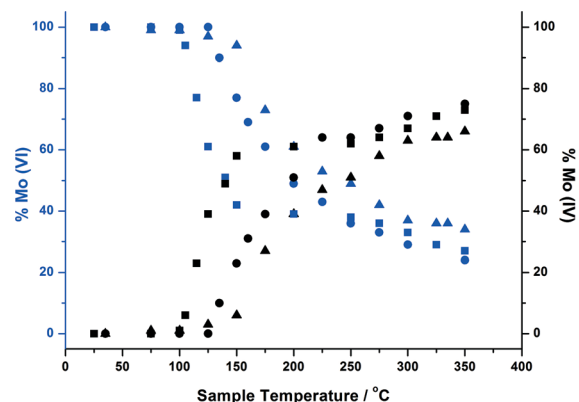


Fig. 9 Oxidation state changes during *in situ* reduction of 1, 3 and 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ ascertained through monitoring the change in absorption edge position from Mo(VI) to Mo(IV) . Reduction studies were carried out in a methanol/He flow whilst heating to 350 °C at a rate of 12 °C min^{-1} . XAFS spectra were collected throughout (~1 min per scan). 1 ML = squares, 3 ML = circles, 6 ML = triangles.

monolayers have a significant contribution from Mo–Mo at 2.6 Å (Table 1). Literature indicates that this could be a bonded distance indicating a Mo^{5+} dimer or Mo^{4+} trimer.³² MoO_2 also contains a Mo–Mo bonding contribution (Fig. 11), however, at a shorter distance of 2.5 Å and with a reduced average co-ordination number.³³ Previous work of Kikutani *et al.*³⁴ has established a 2.6 Å Mo–Mo bond in fixed dimeric active sites. The dark brown colour of the catalyst as opposed to the black colour seen for $\text{Fe}_2(\text{MoO}_4)_3$ has been suggested to be due to an excitation to the energy levels of the 2.6 Å Mo–Mo bonds, a result of mixing of the energy levels in the molybdenum atoms. Simple d-electron transitions of isolated molybdenum would have a much weaker absorbance.

The Mo–M paths are difficult to assign with confidence due to their out of phase nature, however, obvious contribution from metal scatterers is apparent (Table 1). For the 1 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ sample, the required fit was achieved with 2 Mo–Fe contributions at 3.07 and 3.55 Å, which corresponds

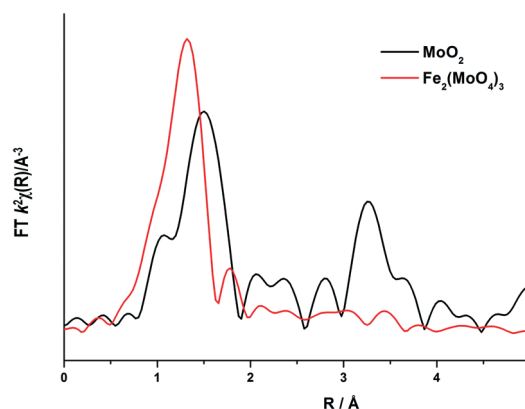


Fig. 10 Magnitude component of the k^2 weighted Fourier transform for the EXAFS data of the 1 and 6 ML $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalyst reduced (procedure as in Fig. 9), and the associated simulated fit performed in Artemis.



Table 1 EXAFS fitting parameters for the 1 and 6 ML MoO_x/Fe₂O₃ catalysts reduced under MeOH/He. Parameters without error values were used as defined input parameters. The R_{factor} for the 1 and 6 ML reduced MoO_x/Fe₂O₃ catalysts were 0.03 and 0.01, respectively

Sample	Abs-Sc	N	R/Å	$2\sigma^2/\text{\AA}^2$	E_f/eV
1 ML MoO _x /Fe ₂ O ₃	Mo–O	2(2)	1.76(5)	0.012(1)	–1(6)
	Mo–O	4.0(1)	2.02(2)	0.003	
	Mo–Mo	1.3	2.68(3)	0.007(2)	
	Mo–Fe	2(2)	3.05(6)	0.007	
	Mo–Fe	2(1)	3.36(2)	0.007	
6 ML MoO _x /Fe ₂ O ₃	Mo–O	1.5(1)	1.78(2)	0.007(7)	0(2)
	Mo–O	3.4(3)	2.03(1)	0.003	
	Mo–Mo	1.3	2.62(2)	0.003(1)	
	Mo–Mo	2.5(1)	2.82(2)	0.01(1)	
	Mo–Mo	4.0(2)	3.37(2)	0.01(1)	

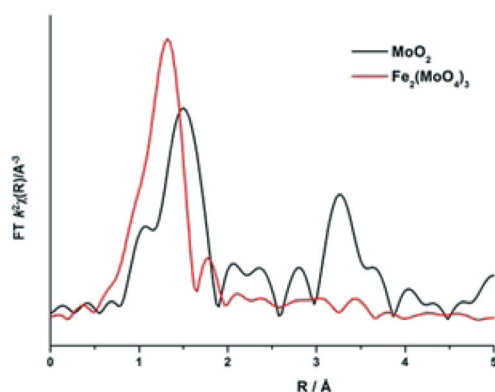


Fig. 11 Magnitude component of the k^2 weighted Fourier Transform data for MoO₃ and Fe₂(MoO₄)₃ (Sigma Aldrich).

well with previous studies of the group, in which we demonstrate the interaction between Mo and Fe.^{7,26} The 3 ML MoO_x/Fe₂O₃ and 6 ML MoO_x/Fe₂O₃ samples are dominated by Mo–Mo contributions, with Mo–Fe distances being difficult to establish.

The EXAFS spectra support the XANES, in which we see a change of structure of the molybdena species under reduction. Fig. S12† and Table S2†, enable comparison to the original catalysts prior to reduction. In studying these ML dosed MoO_x/Fe₂O₃ catalysts, the results are representative of the surface layers. Information obtained through EXAFS alludes to the presence of reduced clusters of Mo after reduction. The reduced molybdenum species may be considered to have molybdenum units with 2.6 Å Mo–Mo bonds, but the extent of aggregation of the Mo units is not clear. Catalytic testing and BET (Fig. S13–15 and Table S3†) have been performed to give an indication of this. TPD was performed on each of the ML MoO_x/Fe₂O₃ catalysts post reduction, along with that of reduced Fe₂O₃ and reduced Fe₂(MoO₄)₃. The spectra all show a strong resemblance to one another, all producing significant amounts of H₂ and CO₂ indicative of formate adsorption. In light of this, it would suggest that there is a large proportion of reduced Fe₂O₃ at the surface (shown to be Fe₃O₄ through XRD (Fig. S9†)), and therefore also significant clustering of Mo. The overall surface area of the catalysts do

not change substantially after reduction (Table S3†), as expected since the catalyst composition is dominated by Fe₂O₃.

The significance of these Mo clusters, is that the possible dimer/trimer unit suggests a multi-centered Mo reaction site, which momentarily forms on reaction with methanol. Current research is limited with regards to the catalytic mechanisms involved in the selective dehydrogenation reaction to methanol. The majority have identified Mo=O sites as possible dehydrogenation centres for the mild oxidation.^{10,35–37} Initial steps involve dissociative adsorption of methanol to yield surface methoxy and hydroxyl species. The subsequent step, the rate-determining step, is commonly considered to be the C–H^{38–43} cleavage of the surface methoxy. The cleaved H will react with the adsorbed hydroxyl to yield water, and consequently the aldehyde will be lost. A study by Allison *et al.*³² uses *ab initio* quantum mechanics to examine the reaction mechanisms of methanol on MoO₃. It was discovered that dioxo sites were essential for methanol activation, with the single site single step proposal unfavorable. It was concluded that the catalytic active site involves two adjacent dioxo units of Mo=O. Chemisorption of methoxy occurs at one of the Mo=O units, which then places a C–H bond in close enough proximity to one of the oxo bonds of the second dioxo unit. Cleavage of the C–H bond occurs with the second unit to yield a hydroxyl, which will likely bridge to an adjacent Mo centre. Succeeding steps involve formaldehyde desorption, with associated H₂O release. Gaseous oxygen initiates surface regeneration to enable the catalyst to continually perform. The (010) surface in MoO₃ has the exact configuration required for these proposed steps, whilst also possessing the lowest free energy, lying parallel to the double layers of MoO₆ octahedra, forming MoO₃. However, later adsorption studies have revealed that a single crystal of MoO₃ favouring the (010) plane, exhibits low saturation coverage. For this reason this face has been excluded from studies. Focus has turned to studying the reaction at the edge facets and defects sites within MoO₃, since here the saturation uptake of methanol correlates well with the number of active centres exposed on non-basal planes. These faces present dual acid-base sites; an unsaturated Mo cation (O–Mo–O) and an adjacent terminal Mo=O.³⁶ Methanol adsorption occurs readily at these bi-functional sites. The oxygen deficient Mo acts as a Lewis acid centre, binding the methoxy,^{40,44} whilst the basic Mo=O serves to abstract a proton.

A more recent paper by Bowker *et al.*⁴⁵ studied the behaviour of bi-cationic oxide catalysts, with specific focus on the mechanism which dominates in the selective oxidation of methanol to formaldehyde. It was discovered that on increasing the amount of Mo added to Fe₂O₃, the product selectivity was determined by the distribution of the dual and single sites of each species. With increased loading, the product detected transitioned from CO under single Mo sites, to formaldehyde with sufficient Mo dosed for dual site activity. This would fit appropriately with our discovered 2.6 Å Mo–Mo distance, supporting the suggestion of a 2 site adsorption centre responsible for the dehydrogenation of methanol.



Conclusions

This work has yielded crucial insights into the structural properties of $\text{MoO}_x/\text{Fe}_2\text{O}_3$ catalysts and by extension to bulk $\text{Fe}_2(\text{MoO}_4)_3$. The *in situ* XAFS investigations detail the evolution the active phase of the catalysis during the pre-catalysis activation of calcination. The activation procedure for the multi-layer systems show that the formation of the T_d $\text{Fe}_2(\text{MoO}_4)_3$ phase is preceded by the formation of MoO_3 , however, an active O_h Mo overlayer remains at all times. These $\text{MoO}_x/\text{Fe}_2\text{O}_3$ systems allow us to probe how the active surface of the catalyst changes under relevant reaction conditions, when coupled with *in situ* XAFS measurements. Indeed, during formaldehyde production no change in Mo speciation is observed although previous studies have identified that surface oxygen is consumed in the reaction.^{7,26} It is therefore apparent that any removal of surface oxygen through extensive reduction of Mo, is followed by rapid re-oxidation by gas phase oxygen. To assess the nature of the Mo when surface oxygen is consumed, *in situ* XAFS measurements in flowing MeOH/He were performed. The reduced state of these systems indicate a significant contribution in the EXAFS data at 2.6 Å, which can be assigned to Mo^{4+} – Mo^{4+} or Mo^{5+} – Mo^{5+} units. From this, it can be inferred that multiple Mo sites are involved in the selective oxidation of methanol to formaldehyde.

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References

- U. R. Pillai and E. Sahle-Demessie, *Appl. Catal., A*, 2004, **276**, 139–144.
- F. Wang and W. Ueda, *Catal. Today*, 2009, **144**, 358–361.
- Y. Meng, T. Wang, S. Chen, Y. Zhao, X. Ma and J. Gong, *Appl. Catal., B*, 2014, **160–161**, 161–172.
- K. Van Der Wiele and P. J. Van Den Berg, in *Comprehensive Chemical Kinetics*, ed. C. H. Bamford and C. F. H. Tipper, Elsevier, 1978, vol. 20, pp. 123–262.
- J. J. Uhlrich, J. Sainio, Y. Lei, D. Edwards, R. Davies, M. Bowker, S. Shaikhutdinov and H. J. Freund, *Surf. Sci.*, 2011, **605**, 1550–1555.
- K. Routray, W. Zhou, C. J. Kiely, W. Gruenert and I. E. Wachs, *J. Catal.*, 2010, **275**, 84–98.
- C. Brookes, P. P. Wells, G. Cibin, N. Dimitratos, W. Jones, D. J. Morgan and M. Bowker, *ACS Catal.*, 2014, **4**, 243–250.
- G. Jin, W. Weng, Z. Lin, N. F. Dummer, S. H. Taylor, C. J. Kiely, J. K. Bartley and G. J. Hutchings, *J. Catal.*, 2012, **296**, 55–64.
- M. Bowker, C. Brookes, A. F. Carley, M. P. House, M. Kosif, G. Sankar, I. Wawata, P. P. Wells and P. Yaseneva, *Phys. Chem. Chem. Phys.*, 2013, **15**, 12056–12067.
- M. Bowker, R. Holroyd, M. House, R. Bracey, C. Bamroongwongdee, M. Shannon and A. Carley, *Top. Catal.*, 2008, **48**, 158–165.
- K. H. Hassan and P. C. H. Mitchell, in *Scientific Bases for the Preparation of Heterogeneous Catalysts: Proceedings of the 10th International Symposium*, ed. E. M. Gaigneaux, M. Devillers, S. Hermans, P. A. Jacobs, J. A. Martens and P. Ruiz, 2010, vol. 175, pp. 475–478.
- B. Ramachandra, J. S. Choi, K. Y. Choo, J. S. Sung, S. D. Song and T. H. Kim, *Catal. Lett.*, 2005, **105**, 23–27.
- M. P. House, A. F. Carley, R. Echeverria-Valda and M. Bowker, *J. Phys. Chem. C*, 2008, **112**, 4333–4341.
- A. P. S. Dias, F. Montemor, M. F. Portela and A. Kiennemann, *J. Mol. Catal. A: Chem.*, 2015, **397**, 93–98.
- A. P. Vieira Soares, M. Farinha Portela and A. Kiennemann, *Catal. Commun.*, 2001, **2**, 159–164.
- M. P. House, A. F. Carley and M. Bowker, *J. Catal.*, 2007, **252**, 88–96.
- N. Pernicone, *Catal. Today*, 1991, **11**(1), 85–91.
- M. P. House, M. D. Shannon and M. Bowker, *Catal. Lett.*, 2008, **122**, 210–213.
- A. M. Beale, S. D. M. Jacques, E. Sacaliuc-Parvaiescu, M. G. O'Brien, P. Barnes and B. M. Weckhuysen, *Appl. Catal., A*, 2009, **363**, 143–152.
- A. M. Beale and G. Sankar, *Nucl. Instrum. Methods Phys. Res., Sect. B*, 2003, **199**, 504–508.
- M. Newville, *J. Synchrotron Radiat.*, 2001, **8**, 322–324.
- B. Ravel and M. Newville, *J. Synchrotron Radiat.*, 2005, **12**, 537–541.
- T. Ressler, O. Timpe, T. Neisius, J. Find, G. Mestl, M. Dieterle and R. Schlögl, *J. Catal.*, 2000, **191**, 75–85.
- M. Massa, R. Haggblad, S. Hansen and A. Andersson, *Appl. Catal., A*, 2011, **408**, 63–72.
- R. Radhakrishnan, C. Reed, S. T. Oyama, M. Seman, J. N. Kondo, K. Domen, Y. Ohminami and K. Asakura, *J. Phys. Chem. B*, 2001, **105**, 8519–8530.
- C. Brookes, P. P. Wells, N. Dimitratos, W. Jones, E. K. Gibson, D. J. Morgan, G. Cibin, C. Nicklin, D. Mora-Fonz and D. O. Scanlon, *J. Phys. Chem. C*, 2014, **118**, 26155–26161.
- S. D. M. Jacques, O. Leynaud, D. Strusevich, A. M. Beale, G. Sankar, C. M. Martin and P. Barnes, *Angew. Chem., Int. Ed.*, 2006, **45**, 445–448.
- M. Bowker, A. F. Carley and M. House, *Catal. Lett.*, 2008, **120**, 34–39.
- M. O'Brien, A. Beale, S. M. Jacques and B. Weckhuysen, *Top. Catal.*, 2009, **52**, 1400–1409.
- A. P. V. Soares, M. F. Portela and A. Kiennemann, *Catal. Rev.: Sci. Eng.*, 2005, **47**, 125–174.



- 31 H. Zhang, J. Shen and X. Ge, *J. Solid State Chem.*, 1995, **117**, 127–135.
- 32 J. N. Allison and W. A. Goddard III, *J. Catal.*, 1985, **92**, 127–135.
- 33 T. Ressler, J. Wienold and R. E. Jentoft, *Solid State Ionics*, 2001, **141–142**, 243–251.
- 34 Y. Kikutani, *J. Mol. Catal. A: Chem.*, 1999, **142**, 247–263.
- 35 J. M. Tatibouët and J. E. Germain, *J. Catal.*, 1981, **72**, 375–378.
- 36 J. M. Tatibouët, *Appl. Catal., A*, 1997, **148**, 213–252.
- 37 M. G. O'Brien, A. M. Beale, S. D. M. Jacques, T. Buslaps, V. Honkimaki and B. M. Weckhuysen, *J. Phys. Chem. C*, 2009, **113**, 4890–4897.
- 38 R. L. Smith and G. S. Rohrer, *J. Catal.*, 1998, **180**, 270–278.
- 39 A. Goodrow and A. T. Bell, *J. Phys. Chem. C*, 2008, **112**, 13204–13214.
- 40 W. L. Holstein and C. J. Machiels, *J. Catal.*, 1996, **162**, 118–124.
- 41 W. E. Farneth, F. Ohuchi, R. H. Staley, U. Chowdhry and A. W. Sleight, *J. Phys. Chem. C*, 1985, **89**, 2493–2497.
- 42 J. S. Chung, R. Miranda and C. O. Bennett, *J. Catal.*, 1988, **114**, 398–410.
- 43 E. M. McCarron and A. W. Sleight, *Polyhedron*, 1986, **5**, 129–139.
- 44 N. Pernicone, F. Lazzerin, G. Liberti and G. Lanzavecchia, *J. Catal.*, 1969, **14**, 293–302.
- 45 A. A. M. Bowker, C. Brookes, E. Gibson, M. House and P. Wells, *Catal. Struct. React.*, 2015, **1**, 95–100.

