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Designing promising ultraviolet (UV) birefringent crystals with different hydrogen-bonded phosphate frameworks†

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In this study, we computationally identified using first-principles calculations two pyridine derivatives with high polarizability anisotropy: (C₆H₇N₂O)⁺ (protonated nicotinamide, abbreviated as 3AP⁺) and (C₆H₆NO₂)⁺ (protonated nicotinic acid, abbreviated as 3CP⁺). Subsequently, we synthesized two novel semiorganic crystals: (3AP) (H₂PO₄) (I) and (3CP) (H₂PO₄) (II). Both crystals incorporate hydrogen-bonded phosphate frameworks (HPFs), which are constituted by H₂PO₄[−] anions linked through O–H...O hydrogen bonds. Compound I (*Fdd2*) features one-dimensional (1D) hydrogen-bonded phosphate frameworks (HPFs) that interconnect with 2D [(3AP)(H₂PO₄)] layers, forming a complex 3D network. In contrast, compound II (*Pbca*) possesses 2D HPFs that are integrated with 3CP⁺ cations to form a 2D layered network. Compound I exhibits a moderate second-harmonic generation (SHG) effect (1 × KH₂PO₄), and a significant birefringence ($\Delta n_{\text{cal.}}$: 0.191@1064 nm). Furthermore, compound II exhibits both a broad bandgap (4.24 eV) but also an exceptional birefringence ($\Delta n_{\text{exp.}}$: 0.284@546 nm), which is the highest value reported among all semiorganic phosphates. This suggests that II could be a promising candidate for use as an outstanding ultraviolet (UV) birefringent material.

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Introduction

Birefringent crystals are key materials used in the production of optical devices, including polarizers, optical isolators, circulators, and phase retarders.^{1–7} Inorganic phosphate crystals, known for their short ultraviolet cutoff edges and good crystal growth habits, represent a classic optoelectronic material.^{8–13} KDP (KH₂PO₄) and KTP (KTiPO₄) are quintessential examples of phosphates, boasting significant application as nonlinear optical (NLO) materials. However, the high symmetry and weak anisotropy of the PO₄^{3−} tetrahedron often limit inorganic phosphates, resulting in low birefringence and phase mismatch.^{14–18} Therefore, the search for phosphates with high birefringence remains an active area of research.

To date, researchers have obtained some metal phosphates with high birefringence ($\Delta n > 0.05$) and proposed several effective strategies, including: (1) introducing lone pair electron cations (such as Sn²⁺, I⁵⁺) or d⁰ transition metal cations (such as Mo⁶⁺, W⁶⁺).^{19–25} These cations can form the building units with high structural distortion and strong polarizability anisotropy, which is beneficial for enhancing the birefringence of phosphates, such as Sn₂PO₄I ($\Delta n_{\text{cal.}}$ = 0.664@546 nm).²¹ (2) Partial substitution of O atoms with X groups (including F, S, or −NH₃) can break the high symmetry of PO₄^{3−} groups, thereby constructing distorted PO_mX_{4−m} tetrahedra with strong polarizability anisotropy, leading to enhanced birefringence, for example, NaNH₄PO₃F·H₂O ($\Delta n_{\text{exp.}}$ = 0.053@589.3 nm),²⁶ KH₂PO₃S ($\Delta n_{\text{cal.}}$ = 0.1@1 μm),²⁷ and NaPO₃NH₃ ($\Delta n_{\text{exp.}}$ = 0.062@546.1 nm).²⁸ (3) Combining PO₄^{3−} with traditional planar π -conjugated birefringent functional groups (such as CO₃^{2−} and BO₃^{3−}). This method can increase the birefringence of compounds ($\Delta n > 0.1$) while also having a wide bandgap ($E_g > 6.2$ eV), such as Sr₃Y(PO₄)(CO₃)₃ ($\Delta n_{\text{cal.}}$ = 0.121@532 nm and E_g = 6.9 eV).²⁹ Notably, Rb[PO₂(NH)₃(CO)₂]·0.5H₂O,³⁰ KPO₂(NHCONH₂)₂,³¹ and NaPO₂(NH)₃(CO)₂³² can be regarded as products designed using both methods (2) and (3) concurrently. Of these, NaPO₂(NH)₃(CO)₂ exhibits very large birefringence and a wide bandgap ($\Delta n_{\text{exp.}}$ = 0.280@550 nm;

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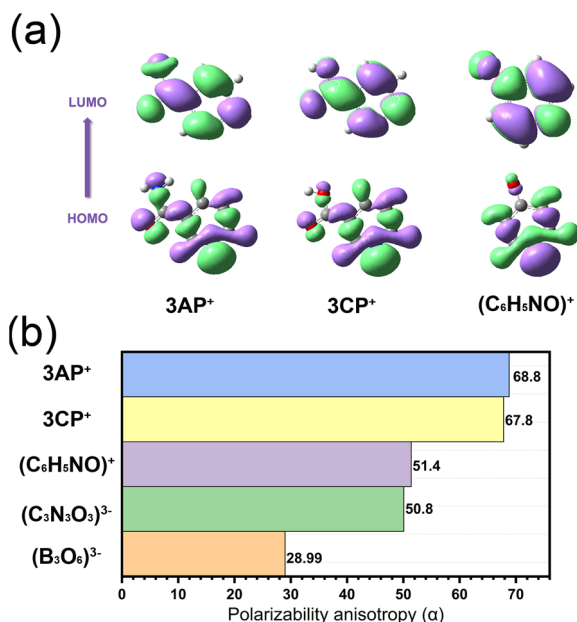
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$E_g \geq 6.5$ eV).³² (4) The “zipper” arrangement of the PO_4^{3-} tetrahedron with large-angle deviation can also lead to the enhancement of birefringence of metal phosphates. For example, $\alpha\text{-YSc}(\text{PO}_4)_2$ exhibits deep ultraviolet transmission and a high birefringence ($\Delta n_{\text{cal.}} = 0.102@1064$ nm).³³

Recently, several semiorganic cations with planar π -conjugated configurations, including $[\text{C}(\text{NH}_2)_3]^+$, $(\text{C}_5\text{H}_6\text{ON})^+$, $(\text{C}_3\text{H}_7\text{N}_6)^+$, *etc.*, have emerged as high-performance anisotropic functional moieties.^{34–37} The combination of π -conjugated organic cations with phosphate tetrahedra has led to the discovery of numerous novel crystals with high birefringence, such as $[\text{C}(\text{NH}_2)_3]_3\text{PO}_4 \cdot 2\text{H}_2\text{O}$ ($\Delta n_{\text{exp.}} = 0.055@546$ nm),³⁸ $[\text{C}(\text{NH}_2)_3]_6(\text{PO}_4)_3 \cdot 3\text{H}_2\text{O}$ ($\Delta n_{\text{exp.}} = 0.078@546$ nm),³⁹ $(\text{C}_3\text{H}_7\text{N}_6)_6(\text{H}_2\text{PO}_4)_4(\text{HPO}_4) \cdot 4\text{H}_2\text{O}$ ($\Delta n_{\text{cal.}} = 0.220@1064$ nm),⁴⁰ and $(\text{C}_5\text{H}_6\text{ON})^+(\text{H}_2\text{PO}_4)^-$ ($\Delta n_{\text{cal.}} = 0.25@1064$ nm).⁴¹ Clearly, as the π -conjugation in organic cations expands, their contribution to birefringence similarly increases significantly. Therefore, we identified two novel birefringent functional units, characterized by higher π -conjugation and stronger polarizability anisotropy, which are $(\text{C}_6\text{H}_6\text{N}_2\text{O})^+$ (protonated 3-carboxamide pyridine, 3AP^+) and $(\text{C}_6\text{H}_6\text{NO}_2)^+$ (protonated 3-carboxypyridine, 3CP^+) (Scheme 1). And then, our efforts have resulted in the discovery of two novel semiorganic phosphates, namely $(3\text{AP})(\text{H}_2\text{PO}_4)$ (**I**) and $(3\text{CP})(\text{H}_2\text{PO}_4)$ (**II**). Experimental characterization and theoretical calculations have revealed that **I** exhibits large birefringence and moderate SHG effect ($\Delta n_{\text{exp.}} = 0.196@546$ nm; $1.0 \times \text{KDP}$). **II** can serve an excellent UV birefringent crystal with the bandgap of 4.20 eV and birefringence of $0.284@546$ nm, the highest value among known semiorganic phosphates.



Scheme 1 The HOMO and LUMO maps (a) as well as polarizability anisotropy (b) of related units.

Experimental

Materials and synthesis

$\text{C}_6\text{H}_6\text{N}_2\text{O}$ (99%) and H_3PO_4 (85%) were used as purchased from Adamas Beta. Crystals of **I** were synthesized using a simple evaporation technique in aqueous solution. The raw reactants of $\text{C}_6\text{H}_6\text{N}_2\text{O}$ (8.2 mmol, 1.00 g) and H_3PO_4 (0.5 mL) were mixed together in deionized water (3 mL) in a glass beaker. The solution was allowed to reach room temperature and then slowly evaporated, resulting in white flaky crystals that precipitated. For the preparation of **II**, the starting materials consisted of $\text{C}_6\text{H}_6\text{N}_2\text{O}$ (2 mmol, 244.25 mg), H_3PO_4 (0.5 mL), and H_2O (1 mL). A mixture of the starting materials was placed into Teflon pouches (23 mL) and sealed inside an autoclave, which was heated to 110 °C for 3 days, and then cooled to 30 °C at a rate of 1.6 °C h^{-1} . Colorless flake-like crystals **I** and **II** were obtained with yields of about 80% and 76% (based on $\text{C}_6\text{H}_6\text{N}_2\text{O}$), respectively.

Single crystal structure determination

Single-crystal X-ray diffraction data for the title compounds were collected using a Rigaku XtaLAB Synergy-DW dual-wavelength CCD diffractometer, equipped with Cu K α radiation ($\lambda = 1.54184$ Å) at 293 K. Data reduction was performed using CrysAlisPro, and absorption correction, based on the multi-scan method, was applied.⁴² The structures of **I** and **II** were determined using direct methods and refined by full-matrix least-squares fitting on F^2 utilizing SHELXL-2014.⁴³ Anisotropic thermal parameters were applied to refine all non-hydrogen atoms. The structure underwent a check for missing symmetry elements using PLATON, and none were found.⁴⁴ Crystallographic data and structural refinements of the compounds are listed in Tables S1–S11.†

Powder X-ray diffraction

Powder X-ray diffraction (PXRD) patterns were recorded on a Rigaku Ultima IV diffractometer with graphite-monochromated Cu K α radiation over a 2θ range of 10° to 70° with a step size of 0.02°.

Thermal analysis

Thermogravimetric analysis (TGA) and differential thermal analysis (DTA) data were obtained, *via* using a Rigaku TG-DTA 8121 unit under an argon (Ar) atmosphere, at a heating rate of 10 °C min^{-1} , ranging from 30 °C to 800 °C.

Optical measurements

Infrared (IR) spectra were recorded on a Thermo Fisher Scientific Nicolet 5700 FT-IR spectrometer using KBr pellets over a range from 4000 to 400 cm^{-1} .

Ultraviolet–visible (UV-vis) spectra, ranging from 200 to 800 nm, were recorded on a PerkinElmer Lambda 750 UV-vis spectrophotometer. Reflectance spectra were converted to an absorption spectrum using the Kubelka–Munk function.⁴⁵

Second harmonic generation measurements

Powder SHG measurements were taken using a Q-switched Nd:YAG laser generating radiation at 1064 nm according to the method of Kurtz and Perry.⁴⁶ Crystalline samples of **I** were sieved into distinct particle-size ranges: 50–70, 70–100, 100–140, 140–200, 200–250, and 250–325 μm . Sieved KDP (KH_2PO_4) samples, within the same particle-size ranges, were used as references.

Birefringence

The optical path difference of the title compounds were characterized using a polarizing microscope (Nikon LV1000) equipped with a Berek compensator.⁴⁷ The average wavelength of the light source was 546 nm. Under orthogonal polarization, the two beams of polarized light pass through the crystal at different speeds, producing interference color phenomena after passing through the analyzer. Polarized light with vibration directions parallel to K_1 , K_2 , K_1' , K_1'' , K_2' , K_2'' represents light with different vibration directions. Polarizers are used to filter out light whose vibration directions are not parallel to the analyzer (AA) or polarizer (PP).

The birefringence (Δn) of microcrystals are calculated with the formula: $R = d \cdot (n_s - n_f) = T \cdot \Delta n$. Herein, R is the optical path difference. The n_s represents the refractive index of the slow light, and n_f represents the refractive index of the fast light. Here, T represents the thickness of the tested single crystal.³⁵

Elemental analysis

The elemental content was measured on Vario EL Cube elemental analyzer from Elementar Analysensysteme GmbH, Germany. The combustion temperature was 800 $^{\circ}\text{C}$.

Energy-dispersive X-ray spectroscopy

Microprobe elemental analyses and elemental distribution maps were measured using a field-emission scanning electron microscope (Phenom LE) equipped with an energy-dispersive X-ray spectrometer (EDS, Phenom LE).

Computational method

The electronic structures and optical properties of the compounds were computed using the plane-wave pseudopotential method within the framework of density functional theory (DFT), as implemented in the total energy code Cambridge Sequential Total Energy Package (CASTEP).^{48,49} For the exchange–correlation functional, we selected the Perdew–Burke–Ernzerhof (PBE) formula within the generalized gradient approximation (GGA).⁵⁰ The interactions between the ionic cores and the valence electrons were modeled using norm-conserving pseudopotentials.⁵¹ The following electrons were considered as valence: C-2s²2p², N-2s²2p³, O-2s²2p⁴, P-3s²3p³, and H-1s¹. The basis set was constructed with plane waves up to a cutoff energy of 750 eV. For both compounds, the self-consistent field (SCF) and optical-property calculations utilized a k -point separation of 0.04 $\text{\AA}^{-1}$ for the numerical integration

over the Brillouin zone. The k -point samplings employed were $4 \times 4 \times 4$ for compound **I** and $2 \times 1 \times 1$ for compound **II**.^{48,49}

To investigate the polarizability anisotropy and electronic structure of selected birefringent units, a systematic computational approach was undertaken using the Gaussian 09 software suite.⁵² The hybrid B3LYP functional was employed at the 6-31G(d,p) level of theory for these calculations. Subsequent analysis of the computational results was performed using the Multiwfn 3.8 software package.⁵³ The polarizability anisotropy was quantified based on the static polarizability values.⁵⁴

Results and discussion

Compound **I**, akin to many previously reported semiorganic phosphates, was synthesized by dissolving a specific stoichiometric ratio of nicotinamide ($\text{C}_6\text{H}_6\text{N}_2\text{O}$) and phosphoric acid (H_3PO_4) in an aqueous solution, followed by slow evaporation. It should be noted that nicotinamide is easily hydrolyzed to nicotinic acid in aqueous solutions (Fig. S1†), a process that can be accelerated by heating.^{55–57} This hydrolysis facilitated the hydrothermal synthesis of compound **II** via the reaction involving nicotinamide and H_3PO_4 at 110 $^{\circ}\text{C}$. The transparent single crystals of **II**, with dimensions of approximately $13 \times 7 \times 3 \text{ mm}^3$, were observed (Fig. 1). Powder X-ray diffraction (PXRD) analysis confirmed the purity of the synthesized samples (Fig. S2†). Field-emission scanning electron microscopy (FESEM) analyses of both **I** and **II** detected the presence of C, N, O and P elements (Fig. S3†). Elemental analysis provided weight ratios of C:H:N in **I** and **II** as 5.97:1.95:8.76 and 5.61:0.99:7.90, respectively, which are congruent with the crystal structure analysis of 6:2:9 and 6:1:8, respectively (Table S12†). This agreement corroborates the presence of $(\text{C}_6\text{H}_7\text{N}_2\text{O})^+$ (3AP^+) and $(\text{C}_6\text{H}_6\text{NO}_2)^+$ (3CP^+) as the organic cations in the structures of **I** and **II**, respectively. The infrared (IR) spectra for **I** and **II** are presented in Fig. S4,† with the detailed assignments of the absorption peaks provided in Table S13,† showing consistency with those of previously reported semiorganic phosphates.^{39–41} Thermal analysis demonstrated that the thermal stability of both compounds decompose above 150 $^{\circ}\text{C}$ (Fig. S5†), comparable to other reported crystals, such as $(\text{C}_5\text{H}_6\text{ON})^+(\text{H}_2\text{PO}_4)^-$ (166 $^{\circ}\text{C}$),⁴¹ $(\text{C}_3\text{H}_7\text{N}_6)_6(\text{H}_2\text{PO}_4)_4(\text{HPO}_4) \cdot 4\text{H}_2\text{O}$ (120 $^{\circ}\text{C}$),⁴⁰ $[\text{C}(\text{NH}_2)_3]_6(\text{PO}_4) \cdot 3\text{H}_2\text{O}$ and (100 $^{\circ}\text{C}$).³⁹

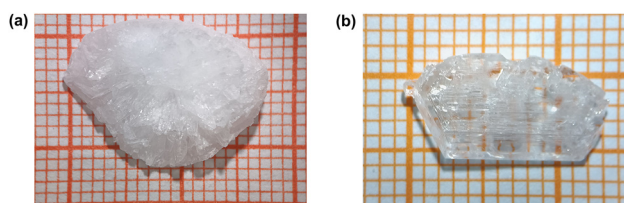


Fig. 1 As-grown crystals of **I** (a) and **II** (b).

Crystal structure

Compound **I** crystallizes in the non-centrosymmetric space group *Fdd2*, with each unit cell housing 16 asymmetric units (Table S1†). Each asymmetric unit consists of one $(\text{C}_6\text{H}_6\text{N}_2\text{O})^+$ (3AP^+) cation and one $(\text{H}_2\text{PO}_4)^-$ anion (Fig. 2a). Within the structure of **I**, each 3AP^+ cation is interconnected with three $(\text{H}_2\text{PO}_4)^-$ anions through three $\text{N}-\text{H}\cdots\text{O}$ and one $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds (Fig. S6a†). Conversely, each $(\text{H}_2\text{PO}_4)^-$ anion is coordinated with three 3AP^+ cations *via* three $\text{N}-\text{H}\cdots\text{O}$ hydrogen bonds and is additionally linked to two axial $(\text{H}_2\text{PO}_4)^-$ anions through two $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds (Fig. S7a†).

Consequently, adjacent $(\text{H}_2\text{PO}_4)^-$ anions are interconnected *via* $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds, resulting in the formation of one-dimensional $[\text{H}_2\text{PO}_4]^-$ chains, denote as 1D hydrogen-bonded phosphate frameworks (1D HPFs) (Fig. 2b). Furthermore, neighboring 3AP^+ cations and $(\text{H}_2\text{PO}_4)^-$ anions are connected through hydrogen bonds, creating a 2D neutral layer of $[(3\text{AP})(\text{H}_2\text{PO}_4)]$ parallel to the *ac* plane (Fig. 2c). Thus, the overall three-dimensional (3D) architecture of **I** can be envisioned as a network of 2D $[(3\text{AP})(\text{H}_2\text{PO}_4)]$ layers interconnected by 1D HPFs (Fig. 2d).

Compound **II** crystallizes in the centrosymmetric space group *Pbca*, with each unit cell comprising 8 asymmetric units

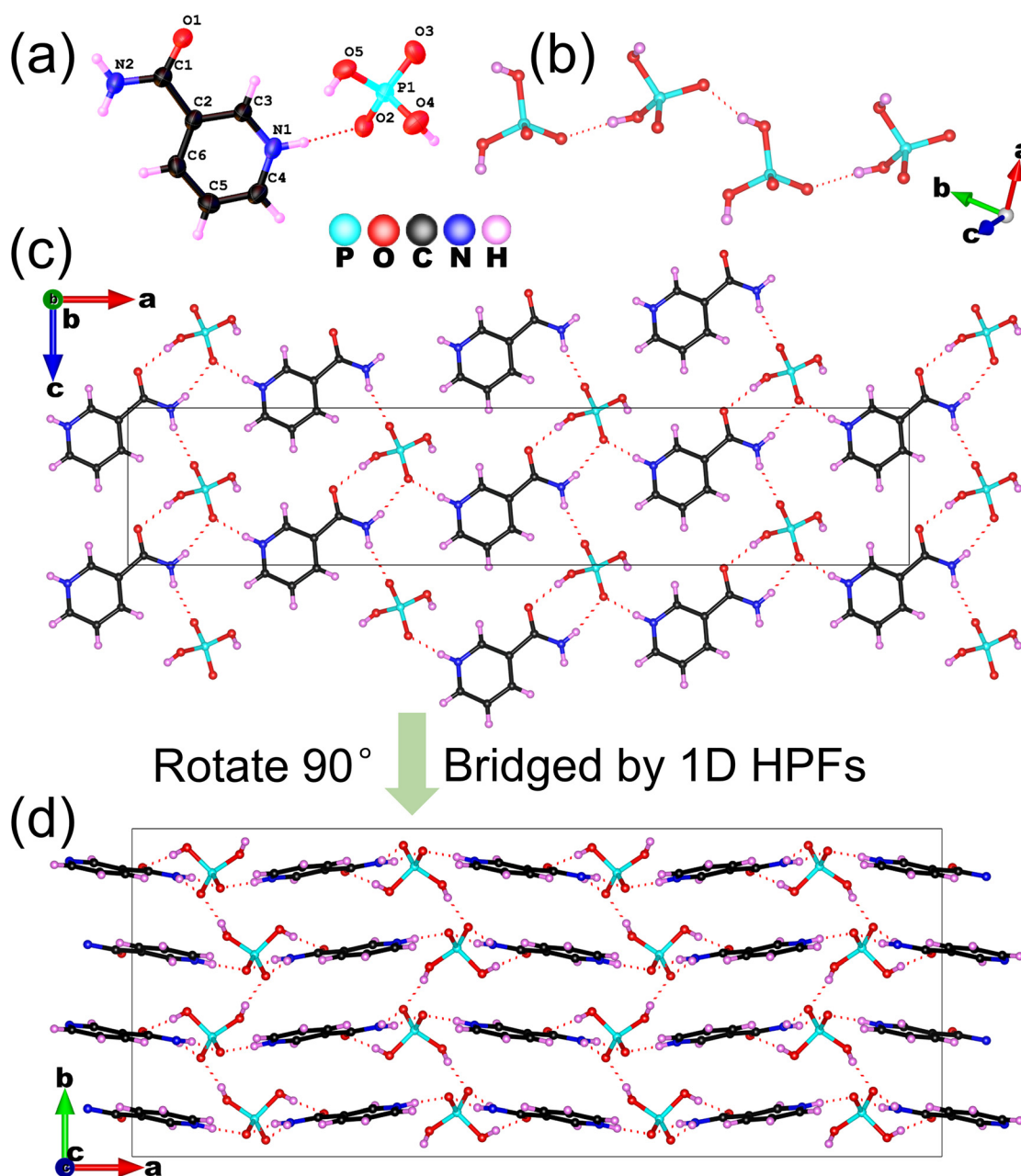


Fig. 2 The asymmetric unit (a), 1D HPFs (b), 2D neutral layer (c) along *b* direction, and 3D overall structure along *c* direction (d).

(Table S1†). Each asymmetric unit consists of one $(\text{C}_6\text{H}_6\text{NO}_2)^+$ (3CP^+) cation and one $(\text{H}_2\text{PO}_4)^-$ anion (Fig. 3a). Within the structure of **II**, pairs of $(\text{H}_2\text{PO}_4)^-$ anions are interconnected *via* $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds, forming a $(\text{H}_4\text{P}_2\text{O}_8)^{2-}$ dimer (Fig. S7b†). These dimers further associate through $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds to create two-dimensional hydrogen-bonded phosphate frameworks (2D HPFs), which are aligned parallel to the *ab* plane (Fig. 3c). Additionally, adjacent 3CP^+ cations are linked into one-dimensional $[\text{C}_6\text{H}_6\text{NO}_2]^+$ chains through weaker $\text{C}-\text{H}\cdots\text{O}$ hydrogen bonds (Fig. 3b and S6b†). These chains are then integrated with the 2D HPFs through $\text{N}-\text{H}\cdots\text{O}$ and $\text{O}-\text{H}\cdots\text{O}$ hydrogen bonds (Fig. S7b†). Consequently, each pair of cationic 1D $[\text{C}_6\text{H}_6\text{NO}_2]^+$ chains is sequentially embedded on either side of the 2D HPFs, generating a neutral 2D $[(3\text{CP})(\text{H}_2\text{PO}_4)]$ layer. These 2D layers interlace with each

other along the *c*-axis, constituting the overall architecture of compound **II** (Fig. 3d).

We conducted a detailed analysis to discern the differences in π - π and dipole-dipole interactions stemming from the dimensional variation of hydrogen-bonded phosphate frameworks (HPFs). In compound **I**, a π - π interaction along the 1D HPFs is evident, characterized by a plane-to-plane distance of 4.5654(15) Å and a dihedral angle of 15.10(13)° (Fig. S8†). Moreover, adjacent 3AP^+ cations bordering the 1D HPFs exhibit larger dihedral angles of 24.295°. In contrast, compound **II** displays a π - π interaction along one side of the 2D HPFs with a plane-to-plane distance of 5.0811(13) Å and a dihedral angle of 9.06(11)° (Fig. S9†). Additionally, neighboring 3CP^+ cations from distinct 2D $[(3\text{CP})(\text{H}_2\text{PO}_4)]$ layers are nearly antiparallel, with a dihedral angle of 0.12(11)° and a plane-to-

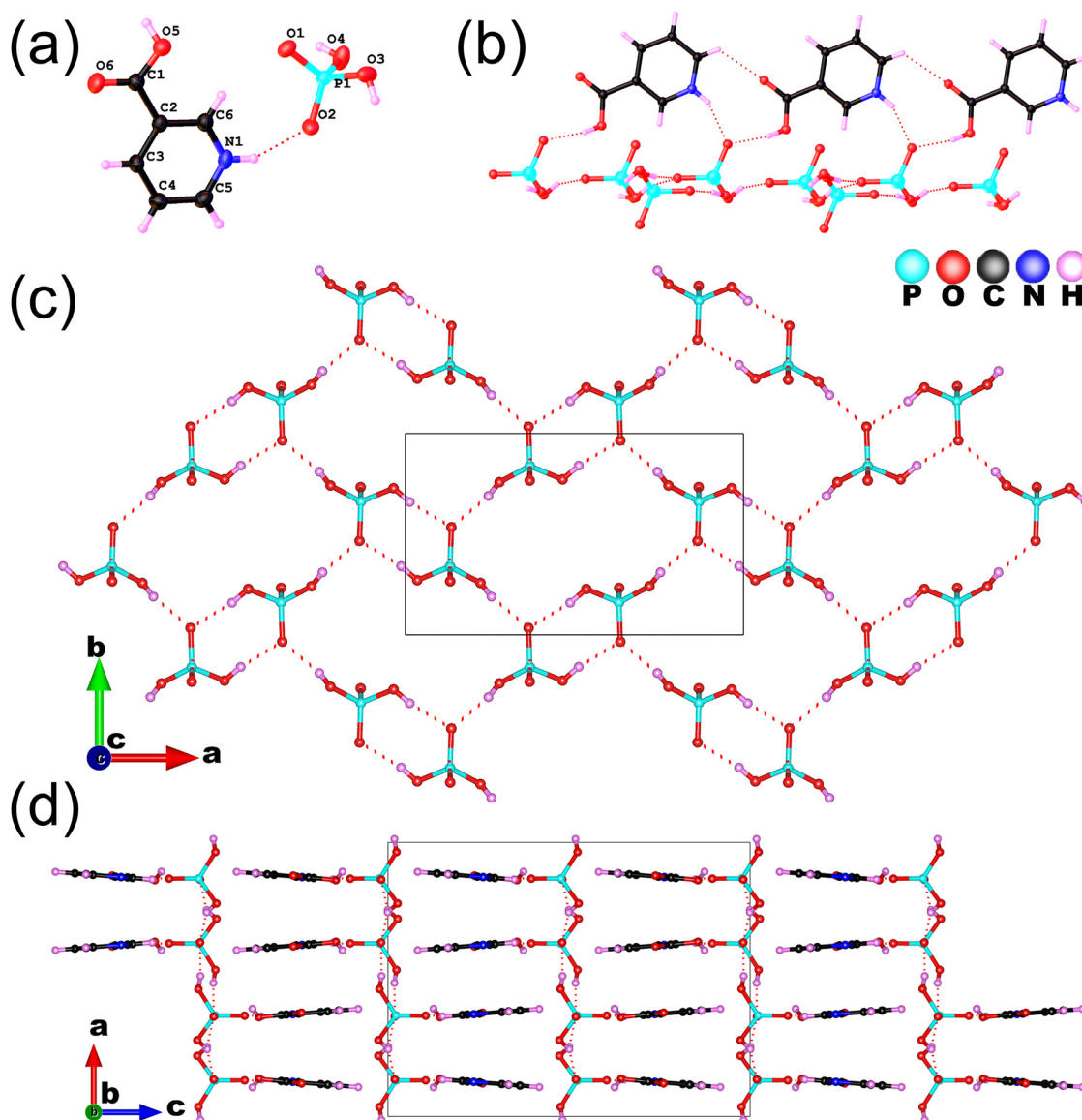


Fig. 3 The asymmetric unit (a), 1D $[\text{C}_6\text{H}_6\text{NO}_2]^+$ chain (b), 2D HPFs (c) along *c* direction, and overall structure along *b* direction (d).

plane distance of 5.1088(13) Å (Fig. S9†). Notably, this π - π interactions act as the binding force for the stacking of the 2D neutral layers in **II**. Consequently, compound **II** demonstrates stronger dipole-dipole interactions, which may contribute to an enhanced birefringence performance.

Transmittance and bandgaps

Ultraviolet-visible (UV-Vis) absorption spectra indicate that compounds **I** and **II** both exhibit bandgaps of 4.24 eV, with cutoff wavelengths at 265 nm and 270 nm, respectively (Fig. 4a and b). Electronic structure calculations substantiate that the two compounds possess indirect and direct bandgaps, with theoretical values of 2.279 eV and 1.919 eV, respectively (Fig. S10†). To mitigate the systematic underestimation inherent in the Generalized Gradient Approximation (GGA) method when calculating optical properties, scissor operators of 1.921 eV and 2.281 eV are applied. Density of states (DOS) analysis reveals that for both compounds, the top of the valence band predominantly originates from O-2p orbitals, while the conduction band minimum is largely composed of C-2p orbitals, with a minor contribution from N-2p and O-2p orbitals (Fig. 4c and d). This suggests that the bandgaps are mainly shaped by the organic constituents, a finding consistent with previous research on semiorganic phosphates.^{39–41}

SHG property

Given that compound **I** crystallizes in a polar space group, we assessed its SHG effect. Upon 1064 nm laser excitation, the

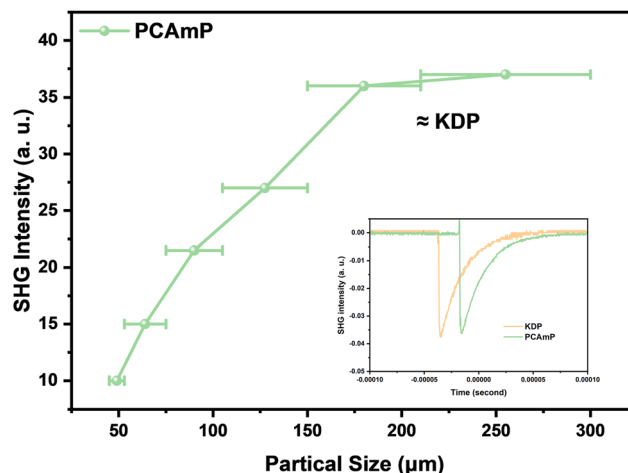


Fig. 5 Phase-matching curve of **I** with 1064 nm laser radiation (insert is oscilloscope traces of the SHG signals for powders of **I** and KDP (150–210 μm) with 1064 nm laser radiation).

SHG intensity of **I** was found to be on par with that of KH_2PO_4 (KDP), utilizing type **I** phase matching (Fig. 5). Compound **I** demonstrates a superior SHG response compared to certain reported semiorganic phosphates, such as $(\text{C}_3\text{H}_7\text{N}_6)_6(\text{H}_2\text{PO}_4)_4(\text{HPO}_4)\cdot 4\text{H}_2\text{O}$ ($0.1 \times \text{KDP}$)⁴⁰ and $(\text{C}_5\text{H}_{12}\text{NO})\text{H}_2\text{PO}_4$ ($0.15 \times \text{KDP}$)⁵⁸ and is comparable to $[\text{C}(\text{NH}_2)_3]_2\text{HPO}_4\cdot\text{H}_2\text{O}$ ($1.2 \times \text{KDP}$)⁵⁹ and $[\text{C}(\text{NH}_2)_3]_2\text{PO}_3\text{F}$ ($1.0 \times \text{KDP}$)⁶⁰. However, it underperforms in SHG effect relative to $[\text{C}$

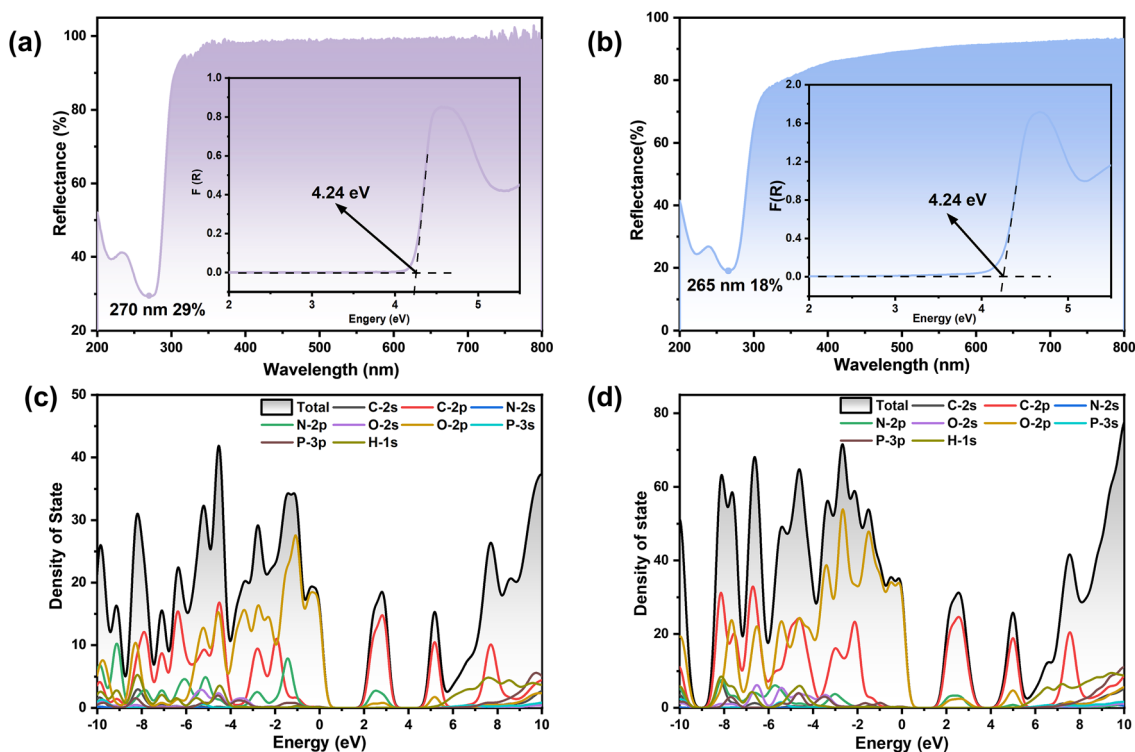


Fig. 4 UV-vis diffuse reflectance spectra (the inset show the bandgaps and $F(R)$ is absorption coefficient/scattering coefficient.) for **I** (a) and **II** (b), as well as the density of states for **I** (c) and **II** (d).

$(\text{NH}_2)_3\text{PO}_4 \cdot 2\text{H}_2\text{O}$ ($1.5 \times \text{KDP}$),³⁸ $[\text{C}(\text{NH}_2)_3]_6(\text{PO}_4) \cdot 3\text{H}_2\text{O}$ ($3.8 \times \text{KDP}$),³⁹ and $(\text{C}_5\text{H}_6\text{ON})^+(\text{H}_2\text{PO}_4)^-$ ($3.0 \times \text{KDP}$).⁴¹ Additionally, theoretical NLO coefficients for compound **I** were calculated, identifying five non-zero independent tensor components: $d_{31} = 0.262 \text{ pm V}^{-1}$, $d_{15} = 0.262 \text{ pm V}^{-1}$, $d_{32} = 0.087 \text{ pm V}^{-1}$, $d_{24} = 0.087 \text{ pm V}^{-1}$, and $d_{33} = -0.426 \text{ pm V}^{-1}$. Since the space group of **II** is *Fdd2*, in the *mm2* point group, its effective tensor should be d_{31} , which is comparable with the measured result.^{61–63}

Birefringence property

Using a polarization microscope equipped with a Berek compensator, we measured the optical path difference (R). This measurement is required to determine the transition of the crystals from orthogonal polarization to complete extinction. By applying the formula $R = \Delta n \times T$, where Δn represents the birefringence and T the sample thickness, we obtained the experimental birefringence values for compounds **I** and **II** (Fig. 6). Using a 546 nm laser source, we observed optical path difference of 1007.41 nm for **I** and 1139.46 nm for **II**, corresponding to sample thicknesses of 5.13 μm and 4.01 μm , respectively (Fig. S11†). The resulting birefringence ($\Delta n_{\text{exp.}}$) were measured as 0.196 at 546 nm for **I** and 0.284 at 546 nm for **II**. The enhanced birefringence of **II** is attributed to the intensified π - π and dipole-dipole interactions stemming from the smaller dihedral angles between neighboring 3CP^+ cations within its structure. Notably, the birefringence exhibited by **I** and **II** significantly exceeds the experimental values of recently reported semiorganic phosphates, such as $[\text{C}(\text{NH}_2)_3]_3\text{PO}_4 \cdot 2\text{H}_2\text{O}$ (0.055@546 nm),³⁸ $[\text{C}(\text{NH}_2)_3]_6(\text{PO}_4) \cdot 3\text{H}_2\text{O}$ (0.078@546 nm),³⁹ $[\text{C}(\text{NH}_2)_3]_2\text{Sb}_3\text{F}_3(\text{HPO}_3)_4$ (0.027@546 nm),⁶⁴ $[\text{C}(\text{NH}_2)_3]\text{SbFPO}_4 \cdot \text{H}_2\text{O}$ (0.151@546 nm),⁶⁴ $\text{NaIn}(\text{C}_2\text{O}_4)(\text{HPO}_4)(\text{H}_2\text{O})_5$ (0.098@546 nm),⁶⁵ $[\text{Te}(\text{C}_6\text{H}_5)_2][\text{PO}_3(\text{OH})]_n$ (0.133@550 nm)⁶⁶

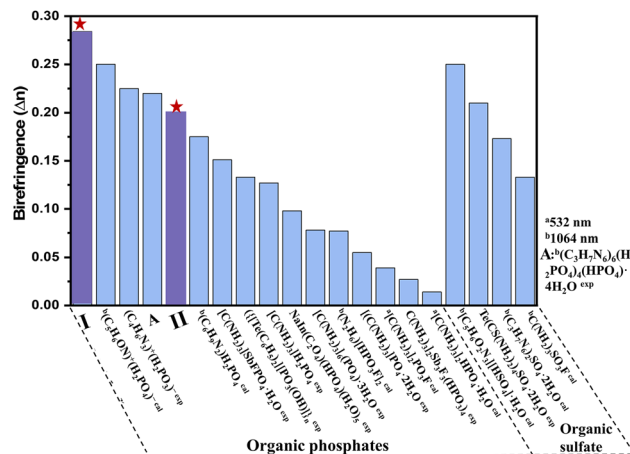


Fig. 7 Birefringence comparison between title compounds and reported phosphates or sulfates containing organic groups.

(Fig. 7 and Table S14†). The increased birefringence of **I** and **II** is primarily due to the enhanced π -conjugation of the 3AP^+ or 3CP^+ cations, which introduces greater anisotropy compared to guanidinium or oxalate. Moreover, the experimental birefringence of **II** surpasses that of commercial birefringent crystals, including $\alpha\text{-BaB}_2\text{O}_4$ (0.122@532 nm), CaCO_3 (0.172@532 nm), and YVO_4 (0.204@532 nm), highlighting its potential as a novel high-performance UV birefringent material.

Subsequently, we computationally determined the dielectric function and refractive indices for compounds **I** and **II**, from which we calculated the theoretical birefringence ($\Delta n_{\text{cal.}}$) using the relationships $\epsilon(\omega) = \epsilon_1(\omega) + i\epsilon_2(\omega)$ and $n^2(\omega) = \epsilon(\omega)$. Both **I** (*Fdd2*) and **II** (*Pbca*) are characterized as biaxial crystals, with their refractive indices ordered as $n_x > n_y > n_z$. Specifically, for

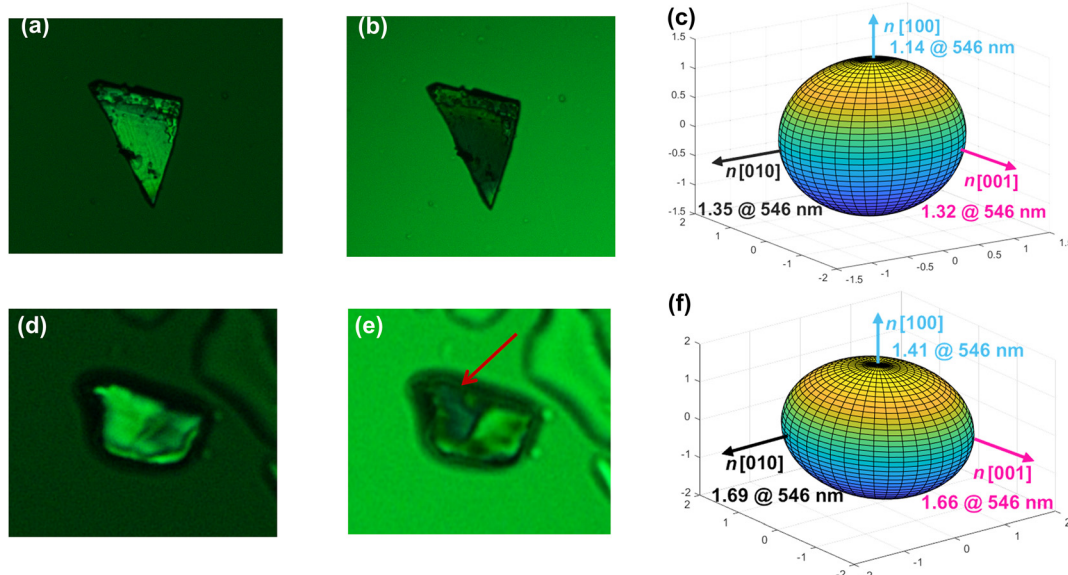


Fig. 6 Original crystal, crystals achieving complete extinction and triaxial ellipsoid of three refractive indices along crystallographic axes for **I** (a–c) and **II** (d–f).

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