

Chiral *N,N'*-dioxide ligands: synthesis, coordination chemistry and asymmetric catalysis†

Cite this: *Org. Chem. Front.*, 2014, **1**, 298

Xiaohua Liu, Lili Lin and Xiaoming Feng*

Received 29th November 2013,
Accepted 5th January 2014

DOI: 10.1039/c3qo00059a

rs.c.li/frontiers-organic

A class of conformationally flexible ligands composed of a tertiary amino oxide–amide backbone and a straight-chain alkyl spacer was developed. These C_2 -symmetric chiral *N,N'*-dioxide ligands could be straightforwardly synthesized from readily available amino acids and amines. They act as neutral tetradentate ligands to bind a wide variety of metal ions. Non-planar *cis*- α $M(N,N'$ -dioxide) complexes enable an intriguing and easily fine-tuned chiral platform for a number of asymmetric reactions. Privileged *N,N'*-dioxide ligands frequently show wide substrate generality and exceptional levels of stereocontrol for a specific catalytic reaction. We describe approaches to the ligand design and synthesis, structure and bonding in coordination complexes, and the recent developments in asymmetric catalysis.

Introduction

Asymmetric catalysis has received considerable attention in the past few years, and it ranks among one of the most important objectives in organic synthesis. Enormous numbers of metal- and organo-catalysts have been developed in producing enantioenriched molecules with maximum efficiency.^{1–3} In particular, the inherent properties of chiral ligand–metal cooperative catalysis make it an attractive platform from which to develop catalytic processes of high reactivity and selectivity.⁴ The main families of successful ligands are greatly aided by several guiding principles. It is widely recognized that conformationally rigid scaffolds and multidentate properties of the ligand play critical roles in allowing it “privilege” in the formation of metal-based catalysts.^{3,5,6} Another frequently used principle in the design of synthetic catalysts is symmetry, because the presence of rotational symmetry in chiral ligands could limit the number of possible isomeric catalyst–substrate complexes.^{7,8} For example, a variety of privileged chiral ligands display C_2 -symmetry and rigid backbones such as bis(phosphine) compounds^{3,9} (e.g., BINAP, Duphos), spirobiindane based ligands,³ and bis(oxazolines).^{3,10} Additionally, with regard to the synthetic criteria, the ligand preparation should be accessible in a modular way and in only a few steps with high yield and optical purity from readily available starting materials. During recent years, our group has explored a new

type of C_2 -symmetric ligands, namely, chiral *N,N'*-dioxides synthesized from cheap optically pure amino acids.^{11,12} These particular ligands have two alkyl amine oxide–amide subunits separated by a straight-chain alkyl spacer. They are full of conformationally flexible subunits¹³ that challenge the notion of the rigid feature of successful ligands (Fig. 1a). Interestingly, *N,N'*-dioxides provide an intriguing arsenal of chiral metal complex catalysts for asymmetric catalysis upon coordination with a variety of metal precursors (Fig. 1). In this paper, we summarize our efforts in the design and synthesis of chiral *N,N'*-dioxide ligands, and the study of their coordination chemistry, as well as their updated applications in asymmetric catalytic transformations.¹¹

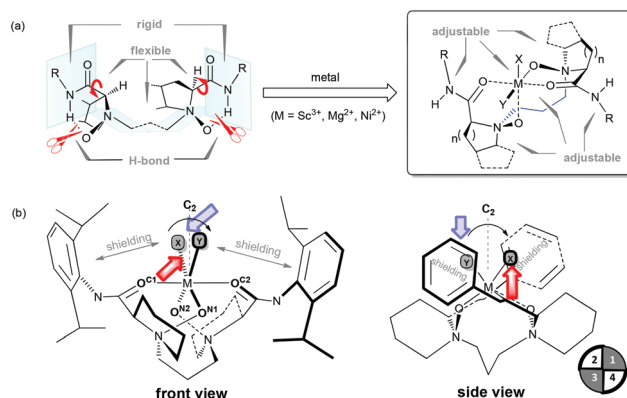


Fig. 1 (a) General structures of chiral *N,N'*-dioxide/metal catalysts. (b) Molecular models of L-metal complex (the isopropyl substituents on the amides are omitted in the side view for clarity).

Key Laboratory of Green Chemistry & Technology, Ministry of Education, College of Chemistry, Sichuan University, Chengdu 610064, China. E-mail: xmfeng@scu.edu.cn
† Electronic supplementary information (ESI) available. See DOI: 10.1039/c3qo00059a

Design and syntheses of *N,N'*-dioxide ligands

N-Oxides are highly polar molecules. The reactions of tertiary amines or pyridine with peracids such as 3-chloroperoxybenzoic acid (*m*CPBA) and hydrogen peroxide easily generate the corresponding *N*-oxides. The basicity and ready availability of *N*-oxides greatly contributes to the abundance of their coordination chemistry.¹⁴ Nevertheless, the application of chiral *N*-oxide-based metal complexes in asymmetric catalysis remained limited to a few successful attempts^{15,16} until we initiated our project. Pyridine *N*-oxides as chiral controllers meet a difficulty of versatility, whereas tertiary amine oxides might suffer from the installation of a new stereocenter at nitrogen. Therefore, to rationally design chiral tertiary amine oxide ligands and catalysts is interesting and challenging.

In 1993, a complete diastereoselective synthesis of homo-chiral *N*-oxides from oxidation of *N*-alkylated prolinamide derivatives was discovered by O'Neil's group.^{17,18} The modular structure offers an opportunity for the design of a new chiral *N*-oxide library. After an initial success in the application of a biquinoline *N,N'*-dioxide¹⁹ and a mono *N*-oxide from L-proline,²⁰ we were intrigued to design *C*₂-symmetric *N,N'*-dioxides in view of the beneficial influence of symmetry and multi-dentate ligands in enantioselective catalysis. *C*₂-Symmetric *N,N'*-dioxides²¹ are readily prepared through straightforward and cost-effective procedures as shown in Scheme 1. Chiral amino acids containing a secondary amino group, such as L-proline, L-pipecolic acid, L-ramipril, and (*S*)-tetrahydroisoquinoline-3-carboxylic acid, could easily be transformed into a number of chiral *N,N'*-dioxides. Take the synthesis of *N,N'*-dioxide **L2** for example.²² Amino amide **3** could easily be obtained from (2*S*,3*aS*,6*aS*)-*N*-Boc-amino acid **1b**. The corresponding *C*₂-symmetric bisamide **4** was generated by coupling of the two amine units *via* 1,3-dibromopropane/*K*₂CO₃ in acetonitrile. It was then treated with *m*CPBA (2.5 equivalents) to give the *N,N'*-dioxide as a single enantiomer without resolution. Further details are given in the ESI.†

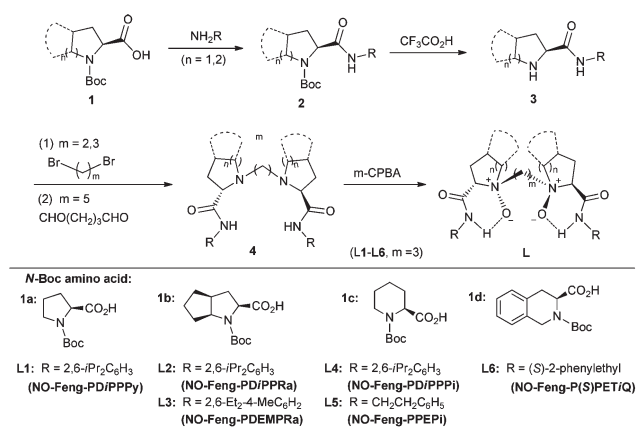
Such a structurally diversified ligand library then leads to the discovery of novel classes of catalysts with very promising

generality. In the early stages of research, *N,N'*-dioxides were used as chiral organocatalysts for Brønsted base and H-bond-assisted reactions.²¹ Nowadays, the utility of this kind of compounds is greatly expanded by using them as privileged ligands for Lewis acid-catalysed asymmetric reactions.¹¹ In light of the potential application of chiral *N,N'*-dioxides, we decide to establish the market using NO-Feng-XXX as a brand name for the whole ligand family (Scheme 1). Both enantiomers of the ligands will be commercially available from J&K Scientific Ltd soon.

Structures of *N,N'*-dioxide-metal complexes

Generating organometallic complexes *in situ* can usually circumvent multiple purifications. When a chiral *N,N'*-dioxide is mixed with an inorganic salt while in an organic solvent at room temperature, a chiral metallic complex system is formed, which is used directly for asymmetric catalysis. Key coordination information can be obtained from the X-ray analysis of crystalline compounds of the *N,N'*-dioxide ligand **L7**¹¹ and the metal complexes^{22–25} (*e.g.*, M = Sc³⁺, Mg²⁺, and Ni²⁺). In the presence of an appropriate metal ion, the intramolecular hydrogen-bonds of the *N,N'*-dioxide disappear, and thereafter, the amide and amine oxide moieties could switch laterally and inwardly to coordinate to the metal center due to the configurational freedom of the sp³ α-carbons¹³ and alkyl spacer (Fig. 1a). *N,N'*-Dioxide acts as a neutral tetradentate ligand that binds a metal ion securely through two amine oxide oxygens and two amide oxygens. The two amides are oriented in an approximately linear fashion pointing in opposite directions. Two coordination sites can be occupied by substrates, solvent molecules or counter anions. Notably, the conformation of the six-membered chelate ring determines an octahedral *cis*-α-*N,N'*-dioxide-metal complex. It differs from the major *trans*- and *cis*-β configuration of the privileged tetradentate salen-metal complexes.^{3,26} The *cis*-α-metal complexes of *N,N'*-dioxides are chiral and expected to be efficient asymmetric catalysts, since there are two coordination sites (X or Y) in a *cis*-position which can capture an ancillary bidentate ligand²³ or two monodentate ligands.^{24,25}

Fig. 1b illustrates the chiral structure created by an *N,N'*-dioxide **L4**-metal complex.²⁴ As shown in the front view, the two amide subunits lie out of the horizontal O^{C1}–M–O^{C2} plane (near a line), front right and rear left, respectively. The two piperidine rings and the alkyl linkage are located below the O^{C1}–M–O^{C2} plane, whereas the above coordination sites ■ hold the positions diagonal to the amine oxide oxygens. Consequently, the two kinds of quadrant of the chiral structure (first and third *versus* second and fourth in the side view) are clearly discriminated spatially, where the first and third ones are sterically crowded by the amides, while the second and fourth ones are partly sterically influenced by the amine oxide moieties, opening for approach of substrates and reagents. It appears that the *cis*-auxiliary ligands ■ can be replaced by at least one of the reagents to give rise to the reacting intermediate, which is suitably oriented to favor a selective attack on one specific face, generating a good level of enantiodifferentiation.



Scheme 1 General synthesis route of *C*₂-symmetric *N,N'*-dioxides and names of selected ligands.

A wide variety of metal cations can associate with chiral *N,N'*-dioxides to form efficient catalysts. The metal list includes rare-earth metals (e.g., Sc^{3+} , Y^{3+} , La^{3+} , Nd^{3+} , Sm^{3+} , Eu^{3+} , Gd^{3+} , Yb^{3+}), transition metals (e.g., Ti^{4+} , Fe^{3+} , Fe^{2+} , Co^{2+} , Ni^{2+} , Cu^{2+} , Cu^+ , Ag^+), and main group metals (e.g., Mg^{2+} , In^{3+}) etc. Therefore, the asymmetric catalysis could also stem from the inherent properties of the metals, such as Lewis acidity, or coordination sphere properties. A difference in the ionic radius of metal ions can be leveraged by the conformational flexibility of the ligands. Table 1 shows some significant structural properties of four metal complexes.^{22–25} Depending on the structures of the *N,N'*-dioxides and the metal precursors, the complexes show distorted octahedral coordinations to some extent. For example, the $\text{X}-\text{O}^{\text{N}1}-\text{O}^{\text{N}2}-\text{Y}$ torsion angles of the **L2**-complexes, with $\text{Sc}(\text{III})$ ²² and $\text{Mg}(\text{II})$ ²⁵ as the chelated cations, are -11.28° and -6.80° , respectively (Table 1). Comparatively, the distortion away from octahedral geometry is much smaller with compound **L4**– $\text{Ni}(\text{II})$ ²³ than the others. It highlights the importance not only of the cation but also of ligand structure in the formation of the chiral messenger in the reactions.

Applications in asymmetric catalysis

A great number of asymmetric catalytic reactions can be catalysed by *N,N'*-dioxide–metal complexes under mild reaction conditions with simple procedures. Several representative examples reported after 2010 have been selected in Table 2 to showcase the diversity of enantioselective reactions realized.²⁷ An early successful example is the hetero-Diels–Alder reaction between Danishefsky-type dienes and a wide range of aldehydes, which is efficiently catalysed by **L4**– $\text{In}(\text{OTf})_3$ complex in high yields, diastereoselectivities, and enantioselectivities.³⁶ It was found that the activity of the pre-prepared catalyst could be maintained even after more than half a year. In some cases, the catalysts were also active without exclusion of water. For example, chiral *N,N'*-dioxide– $\text{Sc}(\text{OTf})_3$ systems could act as efficient catalysts for the asymmetric reduction of ketones with aq. KBH_4 as the reducing reagent,³³ and epoxidation of

Table 2 Selected asymmetric reactions catalysed by *N,N'*-dioxide–metal complexes

L/Metal	Important applications
L1 / $\text{Y}(\text{OTf})_3$ L2 / $\text{Sc}(\text{OTf})_3$	Homologation of α -ketoesters and α -diazoesters ²⁸ Roskamp–Feng reaction; ²⁹ electrophilic addition of α -diazoesters to ketones; ³⁰ ring expansion of isatins with α -diazoesters; ²² Baeyer–Villiger oxidation; ³¹ inverse electron-demand aza-Diels–Alder reaction; ³² reduction reaction ³³
L2 / $\text{Ni}(\text{BF}_4)_2$ L3 / $\text{Mg}(\text{ClO}_4)_2$ L4 / $\text{In}(\text{OTf})_3$ L4 / $\text{Sc}(\text{OTf})_3$	[8 + 2] cycloaddition ³⁴ 1,3-dipolar cycloaddition of nitrile imines ³⁵ Hetero-Diels–Alder reaction ³⁶ α -Functionalization of 3-substituted oxindoles; ³⁷ epoxidation of enones with H_2O_2 ³⁸
L4 / $\text{Ni}(\text{ClO}_4)_2$ L5 / $\text{Sc}(\text{OTf})_3$ L6 / $\text{Sc}(\text{OTf})_3$	Ene reaction ³⁹ Bromoamination ⁴⁰ and iodoamination reaction α -Arylation of 3-substituted oxindoles ⁴¹

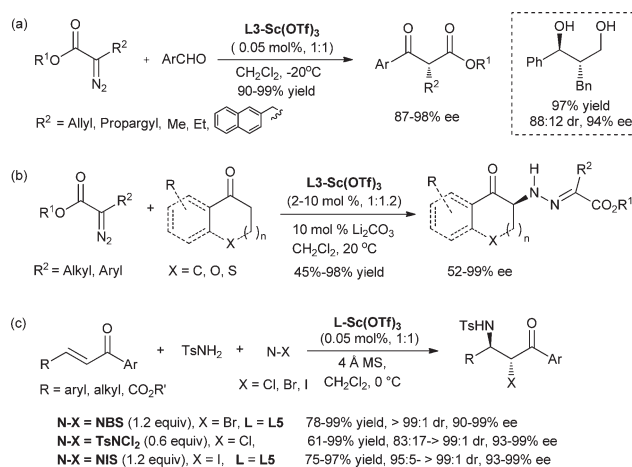
α,β -unsaturated carbonyl compounds using H_2O_2 (0.3%–30% aq.) without additives.³⁸

The general applicability of chiral *N,N'*-dioxide–metal complexes make them useful not only for the practical synthesis of enantiomerically pure compounds but also for the discovery of novel asymmetric processes. Some examples are given to show the particular catalytic characteristics of chiral *N,N'*-dioxide– $\text{Sc}(\text{OTf})_3$ complexes in the control of chemo-, regio- and enantioselectivity. α -Diazocarbonyl compounds have versatile reactivity in organic synthesis.^{22,28–30} Using the highly efficient chiral **L2**– $\text{Sc}(\text{OTf})_3$ catalyst, different reactions of α -diazoesters occur with aldehydes and ketones. Asymmetric Roskamp reaction (also known as Roskamp–Feng reaction)⁴² is a challenging transformation, since it involves the addition of the α -diazoester to an aldehyde and a *H*-shift process in a chemo- and enantioselective fashion to yield chiral α -ketoesters without racemization. A series of α -alkyl- α -diazoesters with aromatic aldehydes could be used to give the chiral α -alkyl- β -ketoesters chemoselectively in 87–98% ee (Scheme 2a).²⁹ The catalyst loading could be reduced to 0.05 mol%. It allowed for convenient separation of the catalyst and the product by flash filtration which thus prevented racemization of the product.

Table 1 Selected geometrical parameters of *M(N,N'*-dioxide)

Metal salt	$\text{Sc}(\text{OTf})_3$	$\text{Sc}(\text{OTf})_3$	$\text{Mg}(\text{OTf})_2$	$\text{Ni}(\text{BF}_4)_2$
<i>N</i> -Oxide	L1	L2	L2	L4
<i>X/Y</i>	$\text{H}_2\text{O}/\text{OTf}^-$	$\text{H}_2\text{O}/\text{OTf}^-$	$\text{H}_2\text{O}/\text{H}_2\text{O}$	*
$\angle 1$ ($^\circ$)	81.54	82.43	85.68	91.13
$\angle 2$ ($^\circ$)	83.17	82.16	85.68	92.40
$\angle 3$ ($^\circ$)	174.35	173.74	176.15	177.35
$\angle 4$ ($^\circ$)	87.89	87.71	81.175	77.79
$\alpha 1$ ($^\circ$)	–12.51	–11.28	–6.80	1.19
$\alpha 2$ ($^\circ$)	90.19	77.09	73.41	103.35
$\text{M}-\text{O}^{\text{N}1}$ (Å)	2.0374	2.0559	2.0096	1.9851
$\text{M}-\text{O}^{\text{N}2}$ (Å)	2.0390	2.0345	2.0153	1.9770
$\text{M}-\text{O}^{\text{C}1}$ (Å)	2.0972	2.0927	2.0640	2.0350
$\text{M}-\text{O}^{\text{C}2}$ (Å)	2.0809	2.1188	2.0443	2.0305
CCDC code	704000	882608	804337	750144

Note: angle: $\angle 1$ ($\text{O}^{\text{N}1}-\text{M}-\text{O}^{\text{C}1}$), $\angle 2$ ($\text{O}^{\text{N}2}-\text{M}-\text{O}^{\text{C}2}$), $\angle 3$ ($\text{O}^{\text{C}1}-\text{M}-\text{O}^{\text{C}2}$), $\angle 4$ ($\text{X}-\text{M}-\text{Y}$); torsion angle: $\alpha 1$ ($\text{X}-\text{O}^{\text{N}1}-\text{O}^{\text{N}2}-\text{Y}$), $\alpha 2$ ($\text{N}^{\text{H}1}-\text{C}^{\text{Ar}1}-\text{C}^{\text{Ar}2}-\text{N}^{\text{H}2}$); *X/Y**: (tetrahydrofuran-2-yl)methanol (Fig. 1b and ref. 22–25).



Scheme 2 Asymmetric catalytic new reactions.

However, an unexpected C–N bond forming reaction of the α -diaoester with ketone takes place with the same catalyst.³⁰ This new reaction involves the electrophilic addition of α -diaoesters to α -alkyl ketones under mild reaction conditions. Optically active alkylhydrazone compounds are obtained in excellent yields (up to 98%) and enantioselectivities (up to 99% ee) for the first time (Scheme 2b). Additionally, **L2**–Sc(OTf)₃ catalysed ring-expansion reaction of α -diaoesters and isatins gives 2-quinolone derivatives,²² whereas succinate derivatives are afforded from **L1**– or **L4**–Y(OTf)₃ complex-promoted homologation of α -ketoesters and α -diaoesters.²⁸ In another notable example, asymmetric haloamination of enones catalysed by **L5**–Sc(OTf)₃ system provided chiral α -halo- β -amino ketones with up to 99% ee and 99:1 dr (Scheme 2c).^{12,40} The issues of the control of both regioselectivity and enantioselectivity of the reaction are well addressed.

Conclusions

In general, the *C*₂-symmetric *N,N'*-dioxide ligands developed by our group form well-organized tetradentate complexes with various metal salts. In a number of asymmetric reactions, *N,N'*-dioxide–metal complexes allow for a remarkable stereochemical outcome under mild reaction conditions. The popularity of privileged *N,N'*-dioxide ligands is due, at least in part, to their flexible configuration, perfectly *C*₂-symmetric, as well as the array of structures accessible from readily available reagents and simple preparation. The exciting results obtained with such a flexible backbone will stimulate further opportunity to develop ideal chiral ligands. There are still unanswered questions concerning the process of the chirality transfer from the chiral *N,N'*-dioxide ligand–metal complex catalyst to the product. More detailed study is anticipated to shed light on the mechanism, which will make the behaviour of the catalyst predictable. We believe that chiral *N,N'*-dioxides will continue contributing to improving the chemistry of asymmetric catalysis.

Conflict of interests

The authors declare no competing financial interest.

Acknowledgements

We thank the National Natural Science Foundation of China (21321061), and the National Basic Research Program of China (973 Program: 2010CB833300) for financial support.

Notes and references

- 1 *Comprehensive Asymmetric Catalysis*, ed. E. N. Jacobsen, A. Pfaltz and H. Yamamoto, Springer, New York, 1999.
- 2 *New Frontiers in Asymmetric Catalysis*, ed. K. Mikami and M. Lautens, John Wiley & Sons, Hoboken, NJ, 2007.
- 3 *Privileged chiral ligands and catalysts*, ed. Q.-L. Zhou, Wiley-VCH, Weinheim, 2011.
- 4 C. A. Sandoval and R. Noyori, in *Organic Chemistry- Breakthroughs and Perspectives*, ed. K.-L. Ding and L.-X. Dai, Wiley-VCH, Weinheim, 2012, ch. 9, pp. 335–363.
- 5 T. P. Yoon and E. N. Jacobsen, *Science*, 2003, **299**, 1691.
- 6 J. H. Xie and Q.-L. Zhou, *Acc. Chem. Res.*, 2008, **41**, 581.
- 7 J. K. Whitesell, *Chem. Rev.*, 1989, **89**, 1581.
- 8 C. Moberg, *Isr. J. Chem.*, 2012, **52**, 653.
- 9 T. Clark and C. Landis, *Tetrahedron: Asymmetry*, 2004, **15**, 2123.
- 10 G. Desimoni, G. Faita and K. A. Jørgensen, *Chem. Rev.*, 2011, **111**, PR284.
- 11 X. H. Liu, L. L. Lin and X. M. Feng, *Acc. Chem. Res.*, 2011, **44**, 574.
- 12 S. X. Huang and K.-L. Ding, *Angew. Chem., Int. Ed.*, 2011, **123**, 7878.
- 13 N. Kumagai and M. Shibasaki, *Angew. Chem., Int. Ed.*, 2013, **52**, 223.
- 14 N. M. Karayannis, *Coord. Chem. Rev.*, 1973, 93.
- 15 G. Chelucci, G. Murineddu and G. A. Pinna, *Tetrahedron: Asymmetry*, 2004, **15**, 1373.
- 16 A. V. Malkov and P. Kočovský, *Eur. J. Org. Chem.*, 2007, 29.
- 17 A. O'Neil, N. D. Miller, J. Peake, J. V. Barkley, C. M. R. Low and S. B. Kalindjian, *Synlett*, 1993, 515.
- 18 A. O'Neil, N. D. Miller, J. V. Barkley, C. M. R. Low and S. B. Kalindjian, *Synlett*, 1995, 617.
- 19 B. Liu, X. M. Feng, F. X. Chen, G. L. Zhang, X. Cui and Y. Z. Jiang, *Synlett*, 2001, 1551.
- 20 Y. C. Shen, X. M. Feng, G. L. Zhang and Y. Z. Jiang, *Synlett*, 2002, 1353.
- 21 Y. H. Wen, X. Huang, J. L. Huang, Y. Xiong, B. Qin and X. M. Feng, *Synlett*, 2005, 2445.
- 22 W. Li, X. H. Liu, X. Y. Hao, Y. F. Cai, L. L. Lin and X. M. Feng, *Angew. Chem., Int. Ed.*, 2012, **51**, 8644.
- 23 K. Zheng, X. H. Liu, J. N. Zhao, Y. Yang, L. L. Lin and X. M. Feng, *Chem. Commun.*, 2010, **46**, 3771.
- 24 Y. L. Liu, D. J. Shang, X. Zhou, X. H. Liu and X. M. Feng, *Chem.-Eur. J.*, 2009, **15**, 2055.
- 25 K. Zheng, C. K. Yin, X. H. Liu, L. L. Lin and X. M. Feng, *Angew. Chem., Int. Ed.*, 2011, **50**, 2573.
- 26 K. Matsumoto, B. Saito and T. Katsuki, *Chem. Commun.*, 2007, 3619.
- 27 The earlier references have been covered in a recent review, see ref. 11. The aim here is to cover the updated developments in our group.
- 28 W. Li, X. H. Liu, F. Tan, X. Y. Hao, J. F. Zheng, L. L. Lin and X. M. Feng, *Angew. Chem., Int. Ed.*, 2013, **52**, 10883.
- 29 W. Li, J. Wang, X. L. Hu, K. Shen, W. T. Wang, Y. Y. Chu, L. L. Lin, X. H. Liu and X. M. Feng, *J. Am. Chem. Soc.*, 2010, **132**, 8532.
- 30 W. Li, X. H. Liu, X. Y. Hao, X. L. Hu, Y. Y. Chu, W. D. Cao, S. Qin, C. W. Hu, L. L. Lin and X. M. Feng, *J. Am. Chem. Soc.*, 2011, **133**, 15268.
- 31 L. Zhou, X. H. Liu, J. Jie, Y. H. Zhang, L. L. Lin and X. M. Feng, *J. Am. Chem. Soc.*, 2012, **134**, 17023.

- 32 M. S. Xie, X. H. Chen, Y. Zhu, B. Gao, L. L. Lin, X. H. Liu and X. M. Feng, *Angew. Chem., Int. Ed.*, 2010, **49**, 3799.
- 33 P. He, X. H. Liu, H. F. Zheng, W. Li, L. L. Lin and X. M. Feng, *Org. Lett.*, 2012, **14**, 5134.
- 34 M. S. Xie, X. H. Liu, X. X. Wu, Y. F. Cai, L. L. Lin and X. M. Feng, *Angew. Chem., Int. Ed.*, 2013, **52**, 5604.
- 35 G. Wang, X. H. Liu, T. Y. Huang, Y. L. Kuang, L. L. Lin and X. M. Feng, *Org. Lett.*, 2013, **15**, 76.
- 36 Z. P. Yu, X. H. Liu, L. Zhou, L. L. Lin and X. M. Feng, *Angew. Chem., Int. Ed.*, 2009, **48**, 5195.
- 37 J. Li, Y. F. Cai, W. L. Chen, X. H. Liu, L. L. Lin and X. M. Feng, *J. Org. Chem.*, 2012, **77**, 9148.
- 38 Y. Y. Chu, X. H. Liu, W. Li, X. L. Hu, L. L. Lin and X. M. Feng, *Chem. Sci.*, 2012, **3**, 1996.
- 39 K. Zheng, X. H. Liu, S. Qin, M. S. Xie, L. L. Lin, C. W. Hu and X. M. Feng, *J. Am. Chem. Soc.*, 2012, **134**, 17564.
- 40 Y. F. Cai, X. H. Liu, Y. H. Hui, J. Jiang, W. T. Wang, W. L. Chen, L. L. Lin and X. M. Feng, *Angew. Chem., Int. Ed.*, 2010, **49**, 6160.
- 41 J. Guo, S. X. Dong, Y. L. Zhang, Y. L. Kuang, X. H. Liu, L. L. Lin and X. M. Feng, *Angew. Chem., Int. Ed.*, 2013, **52**, 10245.
- 42 *Organic Syntheses Based on Name Reactions*, ed. A. Hassner and I. Namboothiri, Elsevier, Oxford, 3rd edn, 2011, p. 408.