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# Surface chemistry in atomic layer deposition of AlN thin films from Al(CH<sub>3</sub>)<sub>3</sub> and NH<sub>3</sub> studied by mass spectrometry†

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Aluminum nitride (AlN) is a semiconductor with a very wide band gap and a potential dielectric material. Deposition of thin AlN films is routinely done by several techniques, including atomic layer deposition (ALD). In this study, we deposited AlN using ALD with trimethylaluminum (TMA) as the Al precursor and ammonia (NH<sub>3</sub>) with and without plasma activation as the N precursor in the temperature range from 100 to 400 °C while monitoring the surface reactions using mass spectrometry. Our results, combined with recent quantum chemical modelling, suggest that the surface chemistry of the deposition process is chemisorption of TMA followed by reductive elimination of the methyl groups to render mono methyl aluminum species. The NH<sub>3</sub> chemisorption is done by ligand exchange to form CH<sub>4</sub> and an –NH<sub>2</sub> terminated surface.

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## 1. Introduction

The properties of aluminum nitride (AlN), including a wide bandgap (6.2 eV), high dielectric constant ( $k \sim 9$ ), high electrical resistivity ( $\rho \sim 10^{11}–10^{13} \Omega \text{ cm}$ ), and high thermal conductivity ( $2.85 \text{ W K}^{-1} \text{ cm}^{-1}$ )<sup>1</sup> make it an important material in microelectronics and optoelectronics. AlN is also used in microelectromechanical systems (MEMS devices) due to its piezoelectric properties.<sup>2</sup> Amorphous AlN finds uses *e.g.*, as passivation and dielectric layers.<sup>3</sup> Thin films of AlN are routinely deposited by reactive sputtering,<sup>4</sup> chemical vapor deposition (CVD),<sup>5</sup> reactive molecular beam epitaxy (MBE),<sup>6</sup> and atomic layer deposition (ALD). ALD of AlN has garnered significant attention for applications demanding robust protective layers, such as the development of corrosion-resistant, insulating and protective coatings.<sup>7</sup>

ALD is based on sequential, self-limiting surface chemical reactions, enabled by carefully tuned deposition chemistry and an alternating precursor supply onto the substrate. This leads to the hallmark of ALD: conformal deposition on topologically complex surfaces. The alternating surface chemistry also facilitates studies of the surface chemistry and reaction mechanisms as the surface is only exposed to one type of precursor molecule at a time. Puurunen *et al.*<sup>8–10</sup> have explored the

surface chemistry of AlN using trimethylaluminum (TMA) and ammonia (NH<sub>3</sub>) by infrared (IR) spectroscopy and nuclear magnetic resonance (NMR) spectroscopy on porous silica and alumina. They found that the amount of hydrogen in amino (–NH<sub>x</sub>) groups increased with the number of reaction cycles on both alumina (Al<sub>2</sub>O<sub>3</sub>) and silica (SiO<sub>2</sub>) substrates, suggesting that more NH<sub>x</sub> groups were formed on the Al<sub>2</sub>O<sub>3</sub> and SiO<sub>2</sub> surfaces as the reactions progressed. The interaction of TMA with dehydroxylated SiO<sub>2</sub> involved ligand exchange with surface hydroxyl (–OH) groups and dissociation in siloxane (–Si–O–Si–) bridges, leading to surface saturation with methyl (CH<sub>3</sub>) groups, followed by the replacement of these CH<sub>3</sub> groups by NH<sub>x</sub> groups upon reaction with NH<sub>3</sub>, indicating a stepwise transformation of the surface chemistry. Additionally, upon reacting with silica, TMA generated –Si–CH<sub>3</sub> groups on the surface, and the subsequent interaction with NH<sub>3</sub>, particularly at higher temperatures, led to the removal of over 90% of the initial CH<sub>3</sub> groups, which aligned with the proposed reaction pathway for this process.

A recent modelling study of the surface reconstruction after the NH<sub>3</sub> pulse in ALD of AlN, GaN and InN,<sup>11</sup> suggests that, at typical ALD temperatures and pressures, the AlN surface is predominantly covered by –NH<sub>2</sub> species at thermal equilibrium. The surface chemistry of TMA and NH<sub>3</sub> in ALD of AlN was also recently mapped out using density functional theory (DFT),<sup>12</sup> suggesting that the decomposition pathways of TMA are crucial for determining film growth rates. Three distinct decomposition pathways for TMA molecules on the surface were observed: (i) TMA undergoes reductive elimination to produce monomethyl aluminum on the surface and ethane

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gas; this pathway does not involve the removal of surface hydrogen atoms, (ii) methyl protonation followed by reductive elimination where TMA interacts with two surface hydrogen atoms, leading to the formation of monomethyl aluminum on the surface, along with methane and ethane gas, surface hydrogen atoms are transferred to the methyl groups during this decomposition pathway, and (iii) methyl protonation twice where TMA reacts with a surface hydrogen atom to generate an aluminum adatom, methane and ethane gas, this process involves the removal of hydrogen atoms from the surface during decomposition.

These pathways were found to be pivotal in determining surface coverage, hydrogen removal, and the overall AlN film-growth efficiency. The DFT study found that TMA would most likely adsorb as a Lewis adduct on the  $-\text{NH}_2$  covered surface and decompose *via* reductive elimination or methyl protonation pathways, with the proximity to previously adsorbed TMA molecules influencing the decomposition pathway. The ratio of the different decomposition pathways was suggested to affect the growth per cycle (GPC) by influencing the saturation coverage of the surface. The relative concentrations of TMA molecules decomposing *via* the three pathways determine the availability of surface sites for TMA adsorption. If the dominant pathway is methyl protonation, which removes surface hydrogen and allows for additional TMA adsorption sites, the GPC may increase due to higher surface saturation coverage. On the other hand, if the dominant pathway is reductive elimination, which does not involve hydrogen removal, the GPC may be lower as the surface may become saturated with hydrogen atoms, limiting additional TMA adsorption sites.

Herein, we use mass spectrometry to investigate the surface reactions in both thermal and plasma ALD of AlN, using TMA as the Al precursor and  $\text{NH}_3$ , with and without plasma activation, as the N co-reactant. The use of *in situ* mass spectrometry for real-time analysis of gas-phase species during the reactions of this precursor combination has not been reported hence it is not fully understood. Our results add to the previous DFT study to form a surface chemical mechanism of ALD of AlN using TMA.

## 2. Experimental details

A hot-wall Picosun R-200 Advanced ALD reactor, equipped with a Litmas remote inductively coupled plasma (ICP) source, was used for deposition. The reactor operated at 4 mbar with a continuous flow of 400 sccm high-purity  $\text{N}_2$  (99.999% with further drying using a getter filter) into the chamber, which was also used as the purge gas. TMA (Epivalence, electronic grade) was kept at 22 °C and vapor was drawn by opening a valve to the low-pressure reactor. The thermal processes used  $\text{NH}_3$  (AGA/Linde, 99.999% and further purified by a getter filter) as the nitrogen precursor whereas the plasma ALD processes used a plasma discharge in a mixture of 75 sccm  $\text{NH}_3$  and 100 sccm Ar (99.999% and further purified by a getter filter). The ICP plasma source was located approximately 75 cm above

the substrate and ignited the plasma with a 2800 W plasma power. Prior to deposition, Si(100) substrates were cleaned with acetone and isopropanol for 10 minutes each to remove surface contaminants before blow-drying them with  $\text{N}_2$  gas.

Film thicknesses were measured by spectroscopic ellipsometry (SE) that uses CompleteEASE software for data analysis and reporting of results. Ellipsometric spectra were recorded in the wavelength range of 210–1690 nm at four angles of incidence ( $45^\circ$ ,  $55^\circ$ ,  $65^\circ$  and  $75^\circ$ ) using a variable angle a J.A. Wollam RC2 Mueller matrix spectroscopic ellipsometer. The thicknesses of the films were then obtained by modeling the measured data using Cauchy dispersion function. Lower film thicknesses (*e.g.*, below 100 nm) were measured by X-ray reflectivity (XRR) with the software PANalytical X'Pert reflectivity and a two-layer model to fit the data and calculate the film thickness. Crystallinity was measured in both  $\theta$ – $2\theta$  and grazing incidence X-ray diffraction (GIXRD) geometries using a PANalytical EMPYREAN MRD diffractometer with a Cu X-ray tube, wavelength of 1.54 Å, operating at 45 kV and 40 mA.

X-ray photoelectron spectroscopy (XPS) was used to obtain elemental compositions and chemical bonding. The results were obtained using Kratos AXIS Ultra DLD, equipped with an  $\text{Ar}^+$  sputtering source. During acquisition, a monochromatic Al  $K\alpha$  X-ray radiation, wavelength 8.35 Å, with 150 W power (anode current = 10 mA and anode voltage = 15 kV) was used. The recorded photoelectron spectra were analyzed using the CasaXPS software package. By quantitative analysis in Casa XPS, the signals originating from the thin films could be de-convoluted and the chemical composition could be quantified. Gaussian–Lorentz functions and Shirley background were used to fit the experimental XPS data.

Optical emission spectroscopy (OES) utilized the Mechelle 900 Spectrometer with a cooled digital 12-bit SensiCam CCD camera system, covering a wavelength range from ultraviolet (UV) to near infrared (NIR), *i.e.*, 200–1100 nm. Positioned perpendicular to the plasma source, the spectrometer captured light emitted from the  $\text{NH}_3$  plasma, directing it to a monochromator through an optical fiber cable. Measurements were conducted immediately after the ICP source in the gas flow direction, before reaching the substrate which is located approximately 65 cm further down which is the only available view port where we can observe the plasma in our reactor set up. Data analysis considered the spectrometer's consistent spectral resolution of 900, enabling the resolution of wavelengths separated by 1/900 of the wavelength.

Gaseous species in the exhaust from the ALD chamber were measured using a Hiden Analytical HPR-30 residual gas analysis (RGA) vacuum process sampling system that includes a HAL 201 RC mass spectrometer (MS) with a Faraday cup detector. The RGA system was connected approximately 80 cm from the substrate stage, with gas lines maintained at room temperature since no residue condensation along the lines was noted. The MS analysis involved collecting both scan spectra and multiple ion detection (MID) of specific ions with electron impact ionization set at 70 eV to ensure efficient reactant molecule fragmentation, enabling the detection and analysis of gas-phase reaction



products. The system was differentially pumped and the operating pressure during typical measurements is kept in the order of  $10^{-6}$  mbar to enable mean free paths that allow collisionless particle transport of the ions from the orifice to the detector. The obtained data was analysed using MASsoft 10 (Hidden Analytical) software.

### 3. Results

Fig. 1 shows how the GPC varies with the TMA pulse time and  $\text{NH}_3$  pulse time: for the plasma (Fig. 1a) and thermal (Fig. 1b) ALD processes. The TMA pulse shows a self-limiting surface chemistry, about 0.2 s pulse is needed to reach saturation in the GPC for both the plasma and thermal processes. The  $\text{NH}_3$  plasma shows a saturating surface chemistry with exposure times higher than 7 seconds (Fig. 1a). The thermal  $\text{NH}_3$  pulse required 12 seconds or longer for self-limiting surface chemistry, albeit less clear than the plasma process (Fig. 1b). It can be noted that the plasma process saturates at 0.83 Å per cycle while the thermal process saturates at approximately 1 Å per cycle.

The plasma process showed a temperature window where the GPC is not affected significantly by temperature between 250 and 325 °C (Fig. 2a). We therefore carried out saturation growth experiments for this process at 300 °C which is within this temperature window. The thermal process showed no clear temperature window where the GPC is not affected by the temperature (Fig. 2b) and very little films were deposited at temperatures below 300 °C. From this, we used 350 °C for studies of the thermal process. Nearly stoichiometric and polycrystalline films were deposited at 300 °C and 350 °C for the plasma and thermal processes respectively, see ESI,† Table S1 and Fig. S1, S2.

The optically active plasma species were studied by OES (Fig. 3), collected when the reactor walls were 300 °C, with a

plasma power of 2800 W, and  $\text{NH}_3$  and Ar flow rates of 100 sccm each. Although the information obtained from OES is limited, relying on excitation processes and radiative transitions, it can give qualitative insight into the species produced in the plasma. The plasma emission lines from OES included NH at 336.1 nm,  $\text{N}_2$  at 357.1 nm, and  $\text{N}_2^+$  at 391.4 nm and 427.6 nm.<sup>13,14</sup> Additionally, the emission peaks at around 336–337 nm indicated the presence of excited NH and  $\text{N}_2$  radicals,<sup>14</sup> corresponding to specific transitions, supporting the existence of  $\text{NH}_x$  and  $\text{N}_2$  species.

A mass spectrometric survey scan was carried out to see the residual reaction species and byproducts leaving the reactor (Fig. 4). The main residual gases from the ALD reactor are  $\text{H}_2$ , N,  $\text{CH}_4$ ,  $\text{NH}_3$ , Ar,  $\text{N}_2$  and  $\text{Al}(\text{CH}_3)$ , while  $\text{CH}_3$  and  $\text{C}_2\text{H}_6$  are detected as minor species. No other species were detected beyond  $m/z = 60$ . From these species,  $\text{CH}_4$  is neither a molecule delivered into the reactor nor a direct fragment of the dosed molecules (TMA and  $\text{NH}_3$ ) but is expected to be a product of surface chemical reactions. Hence, we studied and monitored its evolution with temperature over time as illustrated in Fig. 5. It can be noted that the signal from  $m/z = 16$  increases rapidly with the  $\text{NH}_3$  pulse for both the thermal and the plasma process. As the temperature increases, the  $m/z = 16$  signal is also detected with the TMA pulse. We note that the  $m/z = 16$  signal could also emanate from  $\text{NH}_2^+$ , formed by decomposition of  $\text{NH}_3$  in the QMS. This is likely the explanation for the double peaks observed at the  $\text{NH}_3$  pulses, the first peak is the  $\text{NH}_2^+$  from the  $\text{NH}_3$  pulse, the second is the  $\text{CH}_4^+$  from the surface reactions.

From Fig. 5, activation barriers were calculated using the Arrhenius equation, which relates the rate constant of a reaction to temperature and activation energy. The Arrhenius equation is given by:

$$k = A \cdot e^{-(E_a/RT)}$$

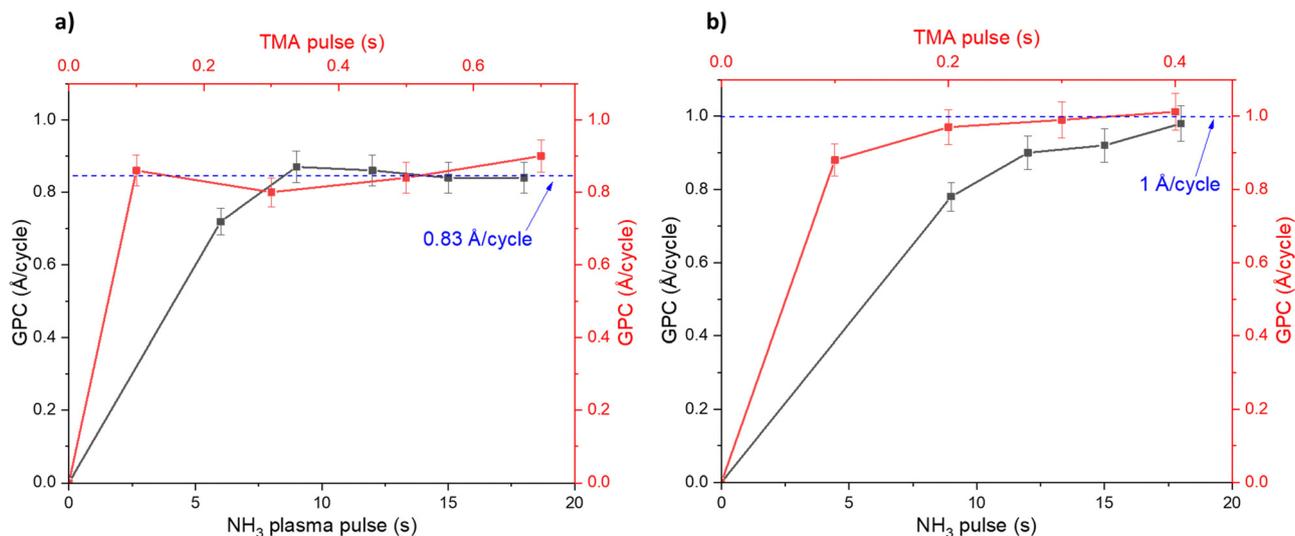


Fig. 1 Saturation curves for the plasma ALD (a) and thermal ALD (b) processes, showing the effect of the TMA pulse time and the (a)  $\text{NH}_3$  plasma time, or the  $\text{NH}_3$  pulse time (b) on the GPC in the ALD processes. The plasma process is studied at 300 °C and the thermal process at 350 °C.



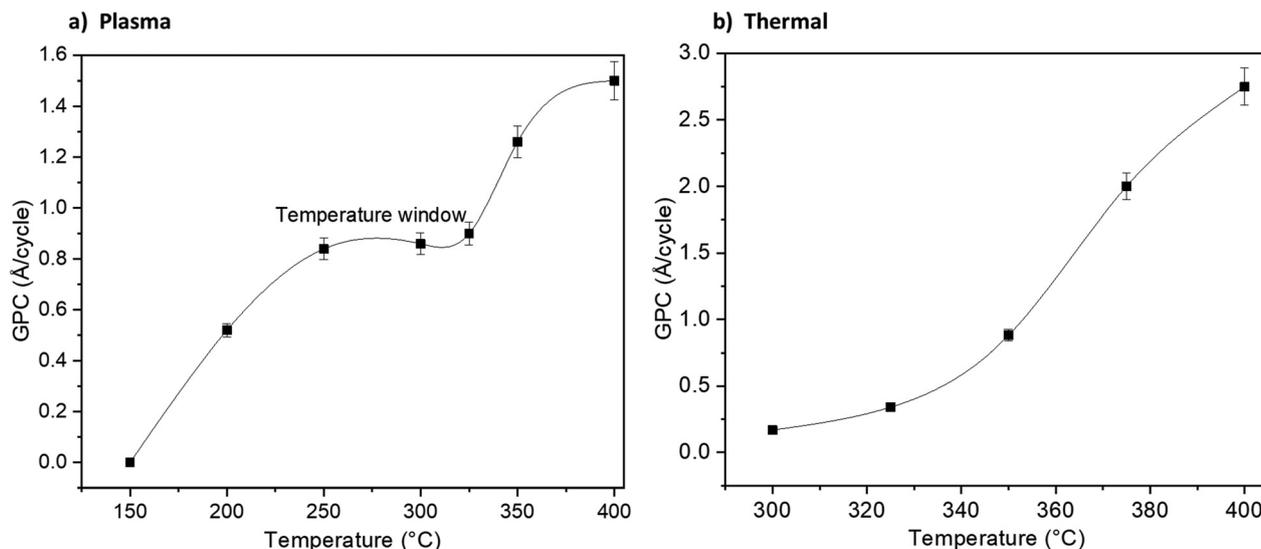


Fig. 2 Growth dependence on process temperature for the plasma (a) and thermal (b) processes using 0.2 s TMA and 12 s  $\text{NH}_3$  pulse times. Lines are a guide for the eyes.

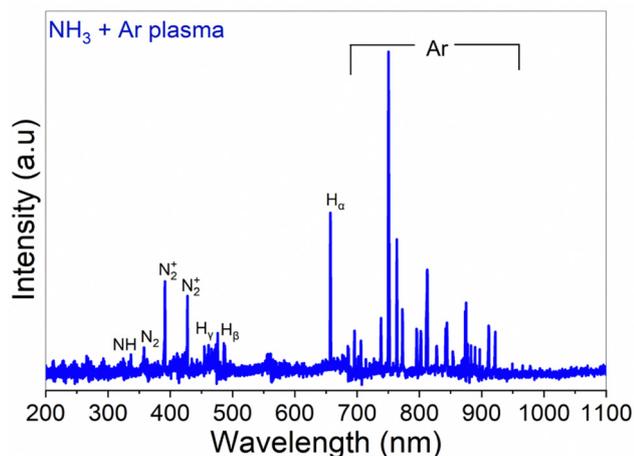


Fig. 3 Optical emission spectra for Ar and  $\text{NH}_3 + \text{Ar}$  plasmas at 2800 W over the wavelength range of 200–1100 nm showing the different reactive species available for reaction with TMA during the plasma process.

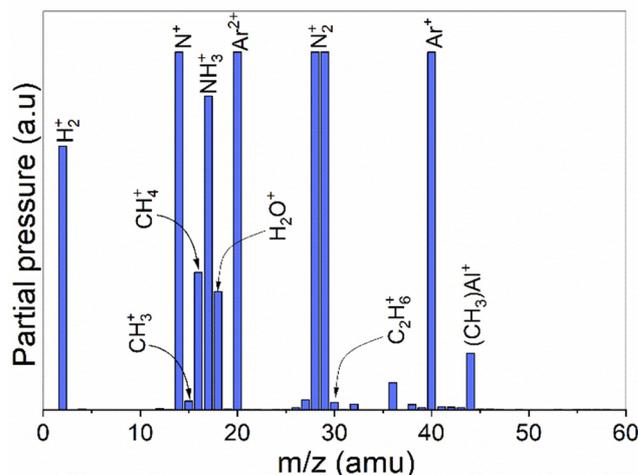


Fig. 4 A survey scan showing average intensities of all detected masses during 5 ALD cycles of the TMA- $\text{NH}_3$  process. No other species were detected beyond  $m/z$  60.

where  $k$  is the rate constant,  $A$  is the pre-exponential factor,  $E_a$  is the activation energy,  $R$  is the gas constant, and  $T$  is the temperature in Kelvin. To extract the activation barriers from experimental data, we performed kinetic measurements at different temperatures, *i.e.*, changes or decay in partial pressure with time, to determine the rate constants. By plotting the natural logarithm of the rate constants ( $\ln k$ ) against the inverse temperature ( $1/T$ ), a linear relationship was obtained, and the activation energy was calculated from the slope of the resulting Arrhenius plot. The pre-exponential factor can be determined from the intercept of the plot as well. Therefore, the activation barriers were extracted by fitting experimental data to the Arrhenius equation and analyzing the resulting plots to determine the activation energy that best describes the temperature dependence of the reaction rate.

From the slope of the tangent of the decay of the  $m/z = 16$  signal, in Fig. 5, we extracted an approximation for the rate constants for the methane evolution from the surface reactions. By doing this for all three studied temperatures, we could approximate the activation energies ( $E_a$ ) for the methane evolution, using Arrhenius plots shown in ESI,† Fig. S3. The extracted activation energies for methane evolution were 19–20  $\text{kJ mol}^{-1}$  during the TMA pulse, 14  $\text{kJ mol}^{-1}$  during the plasma activated  $\text{NH}_3$  pulse, and 17  $\text{kJ mol}^{-1}$  during the thermally activated  $\text{NH}_3$  pulse. The shape of the Arrhenius plots suggests that the  $\text{NH}_3$  pulses are in a mass transport limited regime at  $T \geq 300$  °C, while the TMA pulses seem to be in a surface kinetics limited regime at all temperatures studied. However, considering the small differences in the mentioned activation energies, a comparison between the thermal and



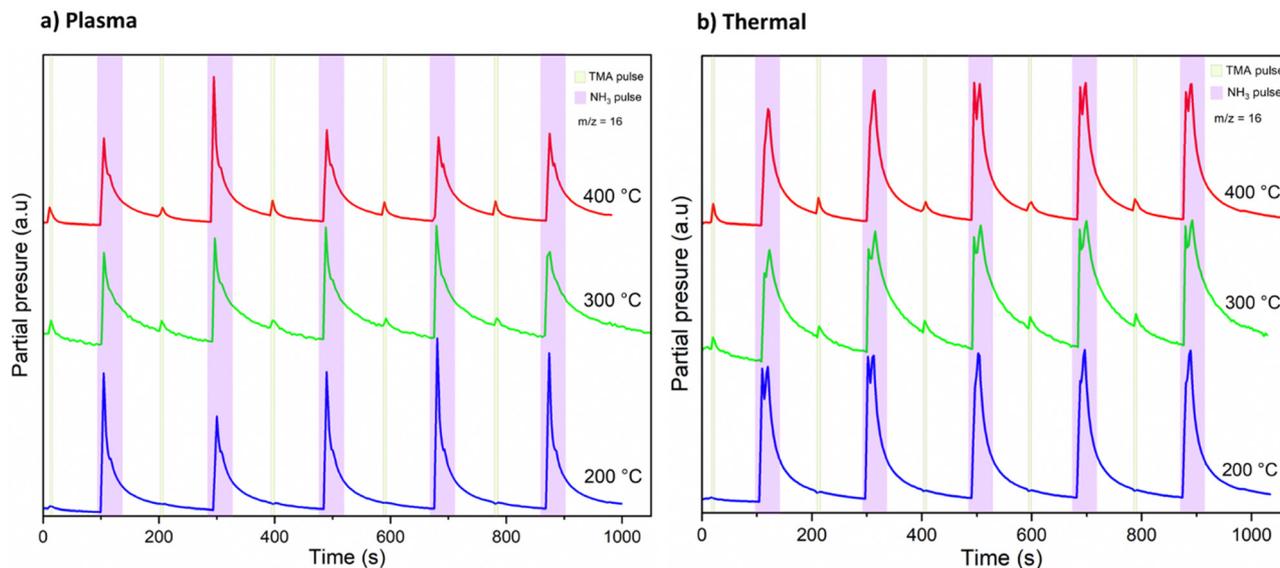


Fig. 5 Mass spectrometric time behaviour of detected  $\text{CH}_4^+$  ( $m/z = 16$ ) during five AlN growth cycles from TMA and  $\text{NH}_3$  in both plasma (a) and thermal (b) processes at different temperatures. The pulse times were fixed at 0.2 s for TMA and 12 s for  $\text{NH}_3$  with 90 s purging in between. The graphs are expressed in arbitrary units and have been vertically shifted for visual clarity.

plasma processes could possibly show that the surface reactions may not be that different.

## 4. Discussion

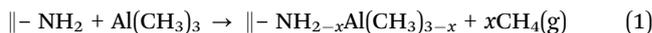
From our observations, we will now discuss the growth mechanisms that were initially proposed in the previous modelling work.<sup>12</sup>

### TMA pulse

DFT suggested that a surface chemistry fully dominated by ligand exchange, render a GPC on the (001) surface of AlN of 2.49 Å per cycle, while a surface chemistry dominated by reductive elimination render a third of that GPC, *i.e.*, 0.83 Å per cycle.<sup>12</sup> Our experiments, with growth on the (100); (002); and (101) AlN surfaces, with increasing TMA and  $\text{NH}_3$  plasma exposure times show a self-limiting surface chemistry for both half reactions (Fig. 1) with 0.83 and  $\sim 1$  Å per cycle at saturation for the plasma and thermal processes, respectively. This suggests that, during the TMA pulse, the surface chemistry is highly dominated by reductive elimination of methyl groups as methane, rendering monomethyl aluminum on the surface. Methane was also detected by mass spectrometry. The change in  $\text{CH}_4$  evolution with temperature, Fig. 5, suggests that thermal activation is stronger from the TMA pulse than the  $\text{NH}_3$  pulse. Our estimated activation energies are in line with this, with the highest activation energy for the methane evolution during the TMA pulse. The much lower intensity of the unprotonated methyl ligand,  $\text{CH}_3$  at  $m/z = 15$  and its possible recombination product  $\text{C}_2\text{H}_6$  at  $m/z = 30$  (two methyls can combine and produce ethane in a ‘termination’ type of reaction), compared to methane at  $m/z = 16$  in Fig. 4, aligns with the anticipated reaction mechanism where the adduct ligand

participates in the reaction, forming  $\text{CH}_4$ , rather than separating independently as  $\cdot\text{CH}_3$ .

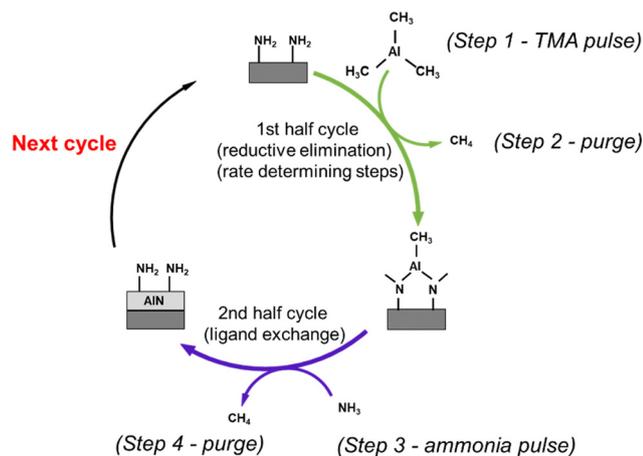
We thus suggest that the TMA molecules adsorb to the surface through their interaction with  $-\text{NH}_2$  surface groups, and this initial half-cycle can be defined by reaction (1):



### $\text{NH}_3$ pulse

Based on the mass spectrometry measurements, where a very sharp  $m/z = 16$  signal was observed during each  $\text{NH}_3$  pulse for both the plasma and thermal processes (Fig. 5), the reaction mechanism for the interaction between  $\text{NH}_3$  and the  $-\text{Al}(\text{CH}_3)_x$  covered surface produces methane. Therefore, we suggest that ammonia molecules, or possibly  $\text{NH}_x$  species in the plasma process, undergo a ligand exchange with surface-bound Al centers to eliminate methane and form an Al–N bond. During the co-reactant half-cycle with  $\text{NH}_3$ , the continued, and even higher intensity compared to that during the TMA pulse, detection of  $\text{CH}_4^+$  ions reaching the mass spectrometer, suggests that more methane is produced in the  $\text{NH}_3$  pulse than in the TMA pulse. This could support the stepwise ligand exchange or methyl protonation mechanism. It is worthwhile to note that during the  $\text{NH}_3$  pulse, especially in the thermal process, some double peaks were observed, which we attribute to either a contribution from  $\text{NH}_2^+$ , formed by decomposition of  $\text{NH}_3$  in the QMS, or instrument artifacts.<sup>15</sup> Instrumental factors such as detector saturation, electronic noise, or ion suppression can possibly affect the shape and resolution of peaks thereby introducing distortions in the mass spectrum, leading to the appearance of forked peaks.



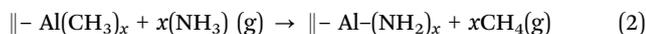


**Fig. 6** An illustration of one ALD cycle for AlN from TMA and  $\text{NH}_3$  showing gaseous TMA reacting with the  $-\text{NH}_2$  groups on the surface, depositing a sub monolayer of Al atoms terminated by  $-\text{CH}_3$  groups followed by  $\text{NH}_3$  which reacts with the  $-\text{CH}_3$  groups converting the surface layer into AlN and recuperating an  $-\text{NH}_2$  terminated surface. Gas phase  $\text{CH}_4$  is released in both the TMA and the  $\text{NH}_3$  pulses (with more  $\text{CH}_4$  signals detected during the second half cycle, rendering the TMA adsorption and its subsequent reaction with  $\text{NH}_2$  groups in the first half cycle the rate-determining steps) as a reaction byproduct that is removed from the reactor by pumping and/or purging before the next ALD half-cycle.

The saturation curve for the plasma  $\text{NH}_3$  pulse (Fig. 1a) shows differences to the thermal  $\text{NH}_3$  pulse (Fig. 1b), suggesting that the thermal reaction is not as strictly self-limited as the plasma reaction. However, such differences are not shown by the residual gas analysis, as both processes show the same pattern of evolution of methane with temperature (Fig. 5). The observation that no reaction occurs below  $300^\circ\text{C}$  in the thermal process agrees with previous studies that showed that the  $\text{NH}_3$  surface reactions with TMA in thermal ALD only occur with reasonable rates at temperatures close to where decomposition of TMA takes place.<sup>16</sup> The presence of hydrogen radicals in the  $\text{NH}_3$  plasma (Fig. 3), suggests that hydrogenation processes of the surface species could play a role in plasma ALD of AlN, which would limit TMA adsorption. This could explain why the plasma process has slightly lower GPCs than the thermal process. The difference in the saturation curves could also be a sign of difference in reaction activity. This is further supported by a higher nucleation delay in the thermal process compared to that in the plasma process (Fig. S4, ESI†).

As the  $\text{NH}_3$  plasma ALD process is expected to be further from thermal equilibrium, the surface termination may not be the same for the plasma process. The detection of  $\text{NH}$  radicals in the OES spectra (Fig. 3) indicates the presence of more reactive nitrogen species in the plasma.

If we assume that both the thermal and plasma process renders  $-\text{NH}_2$  surface species in the  $\text{NH}_3$  pulse, in accordance with recent modelling,<sup>11</sup> the elimination of the  $-\text{CH}_3$  ligands is facilitated by protonation reactions, resulting in the generation of  $\text{CH}_4$  as follows:



A simple growth mechanism between TMA and  $\text{NH}_3$  from our study is illustrated in Fig. 6. The change in the  $m/z = 16$  signal during the TMA pulse with temperature in Fig. 5 suggest that the chemisorption reactions of TMA on the presumably<sup>12</sup>  $-\text{NH}_2$  terminated surface have a significantly higher activation energy than the chemisorption reactions of  $\text{NH}_3$  on the presumably  $-\text{CH}_3$  terminated surface. The  $m/z = 16$  signal during the  $\text{NH}_3$  pulse did not exhibit any temperature dependence in the studied temperature range. This also suggests that the TMA half-cycle is the rate-determining step. The higher levels of C impurities with higher temperature (ESI,† Table S1), are most likely due to TMA decomposition<sup>17</sup> at these temperatures. This suggests that other chemical mechanisms are activated at these temperatures, which is also suggested by the temperature dependence on the GPC in Fig. 2.

## 5. Conclusions

AlN films were grown by both plasma and thermal ALD from TMA and  $\text{NH}_3$ , while the effects of process temperature on precursor reaction mechanisms were studied *in situ* by mass spectrometry. The surface chemical reactions were monitored at  $m/z = 16$ , corresponding to  $\text{CH}_4^+$ , which was produced in both the TMA and  $\text{NH}_3$  pulses. The results from mass spectrometry and the obtained GPC support recent DFT modelling that the reaction pathway for TMA chemisorption on AlN is reductive elimination. From mass spectrometry results, we suggest that the  $\text{NH}_3$  chemisorption pathway involved ligand exchange to produce  $\text{CH}_4$ . Our results also indicated that the reaction activities were much slower for the TMA half-cycle, making it the rate-limiting step.

## Data availability

The data that support the findings of this study are available from the corresponding author upon a reasonable request.

## Conflicts of interest

The authors declare no conflict of interest.

## Acknowledgements

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