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Blue perovskite light-emitting diodes: progress, challenges and future directions

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Metal halide perovskites have excellent optical and electrical properties and can be easily processed via low-cost solution-based techniques like blade-coating and inkjet printing, promising a bright future for various optoelectronic applications. Recently, encouraging progress has been made in perovskite light-emitting diodes (PeLEDs). Green, red, and near-infrared PeLEDs have achieved high external quantum efficiencies of more than 20%. However, as historically blue electroluminescence remains challenging in all previous LED technologies, we are witnessing a similar case with the development of blue PeLEDs, an essential part of displays and solid-state lighting, which lag far behind those of their counterparts. Herein, we review the recent progress of blue PeLEDs and discuss the main challenges including colour instability, poor photoluminescence efficiency and emission quenching by interlayers. Future directions are provided to facilitate the development of efficient blue PeLEDs.

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1 Introduction

Metal halide perovskites with a chemical formula of ABX_3 (A^+ = a monovalent organic or inorganic cation, e.g., $CH_3NH_3^+$

(MA^+) , $CH(NH_2)_2^+$ (FA^+), Cs^+ ; B^{2+} = a divalent metal cation such as Pb^{2+} and Sn^{2+} ; X^- = a halide anion, e.g., Cl^- , Br^- , and I^-) show excellent optoelectronic properties such as a narrow emission linewidth, high absorption coefficient, readily tunable emission wavelengths covering from ultra-violet to near infrared, wide colour gamut, and high charge carrier mobility.^{1,2} In addition, these materials are solution-processable, which can be easily obtained through simple processing from precursor solutions that contain raw materials such as

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AX and BX₂ salts. These outstanding properties of perovskites make them suitable for a wide range of optoelectronic applications including solar cells,³ amplified lasing,⁴ light-emitting diodes (LEDs),⁵ photodetectors,⁶ X-ray detectors,⁷ and so on.

In the past few years, perovskite LEDs (PeLEDs) have shown impulsive progress with external quantum efficiencies (EQEs) boosted from 0.76% to more than 20%, catching up with other LED technologies such as organic and quantum-dot (QD) LEDs.^{8,9} In 2014, Tan *et al.* reported the first PeLEDs that show EQEs of 0.76%, 0.1%, and 0.018% for near-infrared, red, and green emission at room temperature, respectively.¹⁰ Thereafter, tremendous efforts have been devoted to improving the efficiency of PeLEDs through morphology control,¹¹ composition engineering,¹² and device engineering.^{8,13} Very recently, green, red and near-infrared PeLEDs with EQEs of 20.3%,⁸ 21.3%,¹⁴ and 20.7%,¹⁵ respectively, have been reported. These values approach the limits of $\approx 25\%$ for three-dimensional (3D) PeLEDs and $\approx 20\%$ for quasi-two-dimensional (2D) PeLEDs, which are limited by the out-coupling efficiency.¹⁶ The encouraging progress promises a bright future for PeLEDs in the applications of displays and solid-state lighting. However, blue PeLEDs, an essential part of displays and lighting, show limited advancement.

In this review, we summarize the recent progress in blue PeLEDs which are categorized into three sections based on the types of perovskite materials, including 3D perovskites, 2D/quasi-2D perovskites, and colloidal perovskite nanocrystals (PNCs). We discuss the challenges lying in blue PeLEDs and the possible future directions. At present, blue emissive perovskites are mainly achieved by using mixed chloride and bromide, the molar ratio of which can be tuned to adjust the emission colour in the desired blue region. However, mixed-

Cl/Br perovskites show electrical-field-induced phase segregation¹⁷ and reduced photoluminescence quantum efficiency (PLQE). As a result, we will discuss the challenges in terms of colour instability, which is readily observed as a phenomenon of spectral shifts at different voltages or after various time intervals under a constant bias or current, and low PLQE. Plausible strategies are provided to stabilize the colour of blue PeLEDs and improve the PLQEs of blue emissive perovskites. In addition, considering that blue emissive perovskites have wide bandgaps and deep valence band maxima, we will discuss the emission quenching by low-energy interlayers and provide device design strategies for blue PeLEDs. Although we are aware that the operational stability of blue PeLEDs is also significantly important, we do not include it in this review as the operational stability is an open question unsolved for the community of PeLEDs.

2 Progress in blue PeLEDs

2.1 Three-dimensional (3D) perovskites

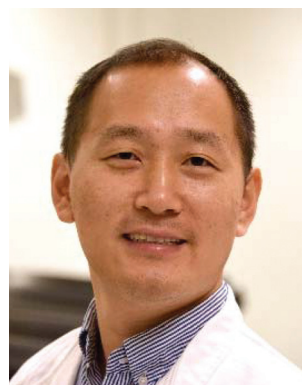
Tunable light emission from sky-blue (490 nm) to deep blue or even ultraviolet (410 nm) can be realized by increasing the chloride content of mixed-Cl/Br perovskites such as MAPb(Br_{1-x}Cl_x)₃ (*x* varies from 0 to 1) (Fig. 1a).¹⁸ The first blue PeLED based on 3D perovskites was reported by Kumawat *et al.* in 2015 using MAPbBr_{1.08}Cl_{1.92}, with a poor brightness of 2 cd m⁻² and an EQE of 0.0001% at room temperature (Fig. 1b and Table 1).¹⁸ Note that the device efficiency is not optimized, which seems to be limited by pinholes in the film morphology and can provide a shunting path in the device.¹⁹ Additionally, PCBM ([6,6]-Phenyl C₆₁ butyric acid methyl ester) used as an electron transport layer (ETL) in the device of ITO/PEDOT:PSS/



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Fig. 1 Bandgap engineering in bulk 3D perovskites and their blue PeLEDs. (a) Tunable blue emissions from 3D-perovskites $\text{MAPb}(\text{Br}_{1-x}\text{Cl}_x)_3$. J - V - L characteristics of (b) $\text{MAPbBr}_{1.08}\text{Cl}_{1.92}$. Reproduced with permission.¹⁸ Copyright © 2015, American Chemical Society and (c) $\text{Cs}_{10}(\text{MA}_{0.17}\text{FA}_{0.83})_{100-x}\text{PbBr}_{1.5}\text{Cl}_{1.5}$ based blue devices; the inset images are the photographs of operational blue PeLEDs. (d) Bias dependent EL spectra of the $\text{Cs}_{10}(\text{MA}_{0.17}\text{FA}_{0.83})_{100-x}\text{PbBr}_{1.5}\text{Cl}_{1.5}$ based blue device. Reproduced with permission.²² Copyright © 2017, John Wiley and Sons.

Table 1 Summary of the reported blue PeLEDs

Perovskite	Device structure	Device performance				Ref.	
		λ_{EL}^a (nm)	L_{max}^b (cd m^{-2})	CIE (x, y) ^c	EQE (%)		
3D	ITO/PEDOT:PSS/MAPbBr _{1.08} Cl _{1.92} /PCBM/Ag	475	2	—	1×10^{-4}	18	
	ITO/Mg:ZnO/MAPb(Br _{1-x} Cl _x) ₃ /CBP/MoO ₃ /Au	450	—	—	—	21	
	ITO/ZnO/Cs ₁₀ (MA _{0.17} FA _{0.83}) _{100-x} PbBr _{1.5} Cl _{1.5} /α-NPD/MoO ₃ /Al	475	3564	0.138, 0.268	1.7	22	
2D/Qusai-2D	ITO/PEDOT:PSS/(PEA) ₂ PbBr ₄ /TPBi/Ca/Al	410	—	—	0.04	25	
	ITO/ZnO/PEIE/(4-PBA) ₂ PbBr ₄ /TFB/MoO ₃ /Al	491	186	—	0.015	26	
	ITO/PEDOT:PSS/PVK/2D-MAPbBr ₃ /TPBi/LiF/Al	456	2	0.166, 0.079	2.4×10^{-3}	34	
	ITO/PEDOT:PSS/POEA ₂ MA _{n-1} Pb _n Br _{3n+1} /TPBi/Ba/Al	480 (494, 508)	19.25	—	1.1	35	
	ITO/PEDOT:PSS/(C ₄ H ₉) ₂ (MA) _{n-1} Pb _n Br _{3n+1} /TmPyPB/LiF/Al	450	1	0.2052, 0.1669	5.4×10^{-3}	36	
	ITO/PEDOT:PSS/(EA) ₂ (MA) _{n-1} Pb _n Br _{3n+1} /TmPyPB/CsF/Al	473 (485)	200	0.16, 0.23	2.6	37	
	ITO/PEDOT:PSS/IPA:PEA ₂ MA:Cs _{n-1} Pb _n Br _{3n+1} /TPBi/LiF/Al	490	2480	0.098, 0.174	1.5	39	
	ITO/PEDOT:PSS/PVK/BA ₂ Cs _{n-1} Pb _n (Br/Cl) _{3n+1} /TPBi/Al	487	3340	—	6.2	38	
	PNCs	ITO/PEDOT:PSS/PVK/CsPb(Br _{1-x} Cl _x) ₃ /TPBi/LiF/Al	455	742	0.157, 0.035	0.07	5
		ITO/PEDOT:PSS/PVK/CsPb(Br _{1-x} Cl _x) ₃ /TPBi/LiF/Al	490	35	—	1.9	52
ITO/ZnO/CsPbBr _{1.5} Cl _{1.5} /TFB/MoO ₃ /Au		480	8.7	—	7×10^{-3}	17	
ITO/PEDOT:PSS/TFB:PFI/CsPb(Br _{1-x} Cl _x) ₃ /TPBi/LiF/Al		469	111	—	0.5	54	
ITO/PEDOT:PSS/TFB:PFI/CsMnPb(Br _{1-x} Cl _x) ₃ /TPBi/LiF/Al		466	245	0.127, 0.077	2.12	57	

^a The peak of the EL spectrum including shoulder peak(s) in the brackets. ^b Maximum luminance. ^c CIE coordinates.

MAPbBr_{1.08}Cl_{1.92}/PCBM/Ag (wherein, ITO is indium-doped tin oxide; PEDOT:PSS is poly(3,4-ethylenedioxythiophene):poly-styrene sulfonate) is a strong emission quenching layer which can reduce the device performance.²⁰

Attempts were made to improve the EQEs of blue 3D PeLEDs through fabricating pinhole-free perovskite films. Lead acetate was used as a lead source in perovskite solutions with ammonium halides, upon which a blue PeLED was fabricated



by using the device structure of ITO/Mg:ZnO/MAPb(Br_{1-x}Cl_x)₃/CBP(4,4'-bis(9-carbazolyl)-1,1'-biphenyl)/MoO₃/Au and demonstrated an EL (electroluminescence) emission peaking at 450 nm at liquid nitrogen temperature (77 K).²¹ Although lead acetate and molar ratio engineering can help to slow down the crystallization process, pinholes still exist. A triple cation strategy with anti-solvent treatment was reported to achieve a pinhole-free and uniform morphology for blue emission.²² A sky-blue PeLED (Fig. 1c), ITO/ZnO/Cs₁₀(MA_{0.17}FA_{0.83})_{100-x}PbBr_{1.5}Cl_{1.5}/NPD(4,4-bis[*N*-(1-naphthyl)-*N*-phenylamino]biphenyl)/MoO₃/Ag, was eventually fabricated and delivered a high brightness of 3567 cd m⁻² and an EQE of 1.7% (Fig. 1c and Table 1).²² The device also shows remarkably stable EL spectra at different bias voltages (Fig. 1d). This suggests that triple cations can help to prevent phase segregation in mixed halide perovskites.²³

2.2 Two-dimensional (2D) and quasi-2D perovskites

2D perovskites, also called layered or Ruddlesden-Popper perovskites (RPPs) with $n = 1$, have a general formula of (RNH₃)₂BX₄, where R is an aryl or alkyl long chain group.²⁴ The first blue PeLED based on 2D perovskites was reported by Liang *et al.* using (PEA)₂PbBr₄ (here, PEA is phenylethylammonium) as the emissive layer, which demonstrated light emission peaking at 410 nm at room temperature with an EQE of 0.04%.²⁵ By changing the organic spacer layer, Cheng *et al.* reported a sky-blue PeLED of ITO/ZnO/PEIE/(4-PBA)₂PbBr₄/TFB/MoO₃/Al (wherein, PEIE is polyethylenimine ethoxylated, TFB is poly(9,9-dioctyl-fluorene-co-*N*-(4-butylphenyl) diphenyl-

amine), and 4-PBA is 4-phenylbutylamine), showing a brightness of 186 cd m⁻² and an EQE of 0.015%.²⁶ The formation of the quantum confinement effect is a very plausible way to realize blue emission from pure halide perovskites; however, 2D perovskites have limited PLQEs due to strong exciton-phonon coupling, which consequently limits the EQEs of blue 2D PeLEDs.²⁷

The combination of 2D and 3D perovskites, named quasi-2D perovskites or RPPs with high n values ($n > 1$), were reported as a useful strategy for promoting the PLQEs.^{28,29} This strategy has been successfully demonstrated in green and infrared PeLEDs with EQEs of more than 10%.^{30,31} In general, quasi-2D perovskites have a formula of (RNH₃)₂A _{$n-1$} B _{n} X _{$3n+1$} , where n is the number of stacking perovskite octahedra. The bandgap can be tuned by changing the n values, becoming attractive as blue emitters of pure halides (Fig. 2a). In addition, these materials not only provide good moisture stability and compact pinhole-free morphology,³² but also have excellent optoelectronic properties such as high PLQEs and good charge carrier balance.³³ Kumar *et al.* prepared quasi-2D perovskites OLA₂MA _{$n-1$} Pb _{n} Br _{$3n+1$} ($n = 1-7$; OLA = oleylammonium) using low- k dielectric organic materials to increase the exciton binding energy by dielectric confinement.³⁴ A blue PeLED with an emission peak at 456 nm was realized, delivering an EQE of 0.0024% and a brightness of 2 cd m⁻². The device performance was further improved by employing a new organic material PEOA (2-phenoxyethylamine), reaching a brightness of 19.5 cd m⁻² and an EQE of 1.1%.³⁵ A similar approach with an excess amount of ligand was used to obtain blue emission



Fig. 2 Band gap engineering in quasi-2D perovskites and their blue PeLEDs. (a) Bandgap tunability of quasi-2D perovskite materials by changing the number (n value) of stacking octahedra. (b) Tunable PL emission by changing the ratio of IPABr to PEA₂(MACs)_{1.5}Pb_{2.5}Br_{8.5} (from 0% to 60%); the inset images are the photographs of deposited quasi-2D perovskite thin films under UV irradiation. (c) EQE-current density plots with emissive layer thickness dependence; the inset image is a photograph of an operational blue PeLED. (d, e) Bias dependent EL spectra. (f) EL spectra recorded after different time intervals. Reproduced with permission.³⁹ Copyright © 2018, Springer Nature.



and achieved a blue PeLED of ITO/PEDOT:PSS/(BA)₂(MA)_{n-1}Pb_nBr_{3n+1}/TmPyPB/LiF/Al (wherein, BA is *n*-butylammonium; TmPyPB is 1,3,5-tri(*m*-pyridin-3-ylphenyl) benzene) with 0.0054% EQE.³⁶ Wang *et al.* used ethylammonium bromide (EABr) to fabricate quasi-2D perovskite thin films and realized a brightness of 100 cd m⁻² and an EQE of 2.6% for blue emission peaking at 473 nm.³⁷ Very recently, a high-efficiency blue PeLED was achieved based on quasi-2D perovskites, showing an EQE of 6.2% with an emission peak of 487 nm.³⁸

Although quasi-2D perovskites have a formula with a constant *n* value, the resulting films prepared following the nominal formula are usually a mixture of RPPs with various *n* values.^{34,35} As a result, quasi-2D perovskites can show double or triple peaks in EL spectra, which could lead to colour instability under the operation of blue devices.³⁹ However, a careful selection of organic spacer materials can stabilize the EL spectra. For instance, Xing *et al.* used PEABr, IPABr (iso-propylammonium bromide), PbBr₂, CsBr and MABr to synthesize quasi-2D perovskites for blue emission (Fig. 2a and b), realizing a blue emissive perovskite with a high PLQE of 88% the PL spectrum of which peaks at 477 nm.³⁹ Finally, they reported a sky-blue PeLED peaking at 490 nm with a brightness of 2480 cd m⁻² and an EQE of 1.5% (Fig. 2c). In addition, the colour stability was tested by recording the EL spectra at different driving voltages from 5 to 13 V and after various time intervals. As shown in Fig. 2d–f, the shape of the EL spectra remains the same at different driving voltages and under different time intervals, suggesting excellent colour stability of the quasi-2D blue device.

2.3 Colloidal perovskite nanocrystals (PNCs)

2.3.1 Synthesis of PNCs

Hot-injection method. The hot-injection method is one of the pioneering techniques for synthesizing CdSe nanocrystals. It also has been widely used for synthesizing other nanocrystals

including PNCs. Fig. 3a (left panel) shows a schematic diagram of the hot-injection setup. In general, lead halides are dissolved in 1-octadecene (ODE) together with organic ligands (*e.g.*, oleylamine and oleic acid). The solution is then stirred at a certain temperature in nitrogen during which Cs-oleate (Cs-OA) solution is added.¹ The mixed solution is then cooled down in an ice bath. The precipitates are collected and purified with nonpolar solvents such as toluene and hexane. A colloidal PNC solution can be obtained when the precipitates are dispersed into a nonpolar solvent. Success in synthesizing blue emitting colloidal PNCs using the hot-injection method was achieved by Protesescu and co-workers, when they obtained a series of highly luminescent colloidal CsPb(Br_{1-x}Cl_x)₃ (0 < *x* < 1) PNC solutions with PLQEs up to 90%.¹ Fig. 3b shows the photographs of the colloidal PNC solutions under UV light and their PL spectra, the emission colour of which was tuned by changing the bromide to chloride ratio. In addition, reducing the size of CsPbBr₃ PNCs from 11.8 nm to 3.8 nm can also tune the emission wavelength from 510 nm to 450 nm because of the quantum confinement effect.¹ Note that the properties of PNCs are sensitive to the reaction temperature and the chain length of the organic ligands, which can control the size of PNCs.^{40–42} But the reaction time has a negligible influence as the formation of PNCs happens within several seconds.⁴³

Ligand-assisted reprecipitation (LARP) technique. The LARP method is simple and versatile compared to the hot-injection method. Fig. 3a (right panel) shows the setup for LARP, where the perovskite precursor solution containing organic ligands is added dropwise into a nonpolar solvent.⁴⁴ Immediately, colloidal PNC solution can be obtained. However, continuous addition of the precursor solution will lead to the formation of precipitates, which may not be dispersed into nonpolar solvents.

Although both the hot-injection and LARP methods can generate blue-emitting colloidal PNCs, the organic ligands

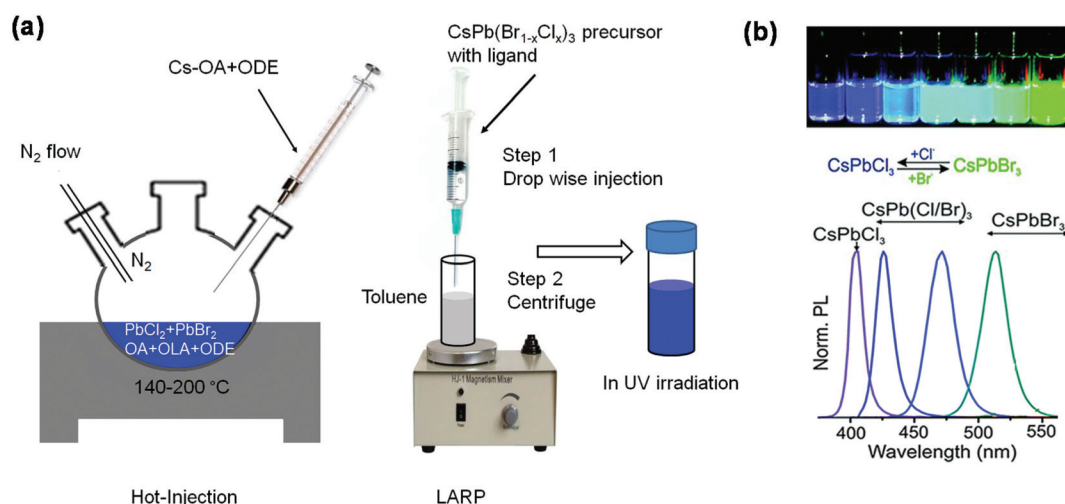


Fig. 3 Synthesis of colloidal PNCs and bandgap tuning *via* anion-exchange. Schematic diagrams of (a) the hot-injection method (left panel) and the LARP technique (right panel). (b) Photographs of colloidal CsPb(Br_{1-x}Cl_x)₃ PNC solutions under UV irradiation and their PL spectra, wherein emission shifts towards the deep blue region with higher chloride/bromide ratios. Reproduced with permission.¹ Copyright © 2015 American Chemical Society.



adhered to PNC surfaces play an important role in their stability and efficiency.⁴⁵ The amount of organic ligands can be adjusted to control the nucleation and growth of PNCs. The shape, size and surface properties of PNCs also depend on the surface ligands other than the processing conditions.⁴⁰ For instance, Sun *et al.* used oleic acid and oleylamine as capping agents and obtained monodisperse cubic shape PNCs; however, excess oleylamine can change the shape of the PNCs from cubic to nanorods and nanowires.⁴⁶

In addition, PNCs with different shapes can be realized by varying the chain length of the amine and acidic group of the ligands.⁴⁷ Moreover, it was reported that the modification of the organic ligands can change the stability and PLQEs of PNCs. Colloidal PNCs synthesized with oleic acid and oleylamine ligands are not stable in storage because of the dynamic change in the binding with the surface of PNCs. An amine-free strategy was reported to improve the stability of the colloidal solution using oleic acid only.⁴⁸ Bulky ligands with steric hindrance such as POSS (polyhedral oligomeric silsesquioxane), TPTES ((3-aminopropyl)triethoxysilane), and TOPO (tri-octylphosphine oxide), are also conducive to slow down the degradation process and thus improve the stability of PNCs.^{49–51} These ligands can also passivate the surface defects of PNCs and therefore increase their PLQEs up to 55%.⁵² However, organic ligands, especially those with long alkyl chains, have poor electrical conductivity, which will form carrier injection barriers in LEDs and thus will lead to high driving voltages and poor device performance.⁵ A balance between the stability of the colloidal solution of PNCs and their electrical properties

is a crucial gap and requires attention from the perovskite materials synthesis community.

2.3.2 Blue PeLEDs based on colloidal PNCs. The first PNC-based PeLED was reported by Schmidt *et al.*, in which MAPbBr₃ PNCs with good stability and high PLQEs were used.² However, the EQE was poor because of the poor morphology of the thin film deposited from the colloidal PNC solution. After that, Xing *et al.* synthesized MAPbBr₃ PNCs using the LARP method and fabricated green LEDs with EQEs up to 3.8%.⁵³ Using a similar approach, Song *et al.* reported efficient CsPbBr₃ PNCs the sub-bandgap states of which were passivated by using ZnBr₂ and FA, and realized green LEDs with a high EQE of 16.48%.⁴⁵ Using an anion-exchange approach, Chiba *et al.* reported 21.3% EQE for red PNC-based PeLEDs.¹⁴ However, the development of blue PNC-based PeLEDs lags far behind those of green and red counterparts.^{8,14,45}

The first attempt on blue PNC-based PeLEDs was made by Song *et al.*, who synthesized blue CsPb(Br_{1-x}Cl_x)₃ (0 < x < 1) PNCs using the hot-injection method, and fabricated a blue PeLED with a turn-on voltage of 5.1 V (at 1 cd m⁻²), an EQE of 0.07%, and a maximum brightness of 742 cd m⁻².⁵ Fig. 4a shows a high-resolution transmission electron microscopy (HRTEM) image of blue emissive CsPbBr_{1.2}Cl_{0.8} PNCs, which have a cubic shape with an average size of 8 nm. The PL (dotted line) and EL (solid line) spectra of these PNCs are shown in Fig. 4b, a peak of which at about 455 nm with 20 nm fwhm (full width at half maxima) can be observed. Fig. 4c demonstrates the device structure; they used poly(9-vinylcarbazole) (PVK) as the hole transport layer (HTL) which can also

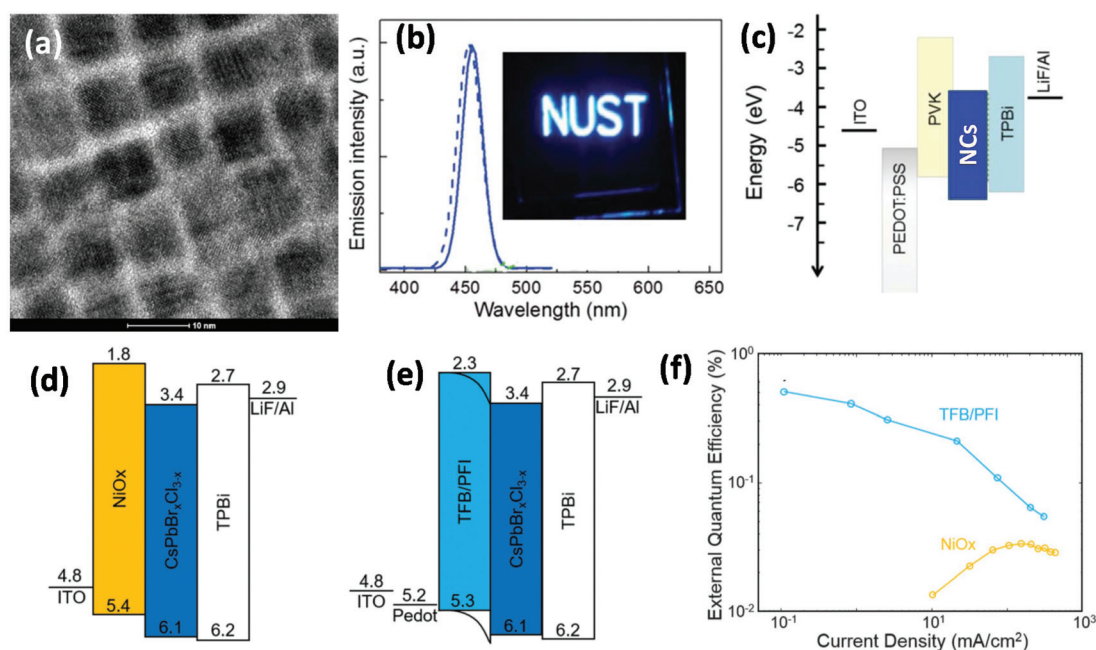


Fig. 4 Blue PNC-based PeLEDs. (a) HRTEM image with a scale bar of 10 nm, and (b) EL (solid line) and PL (dotted line) spectra of CsPbBr_{1.2}Cl_{0.8} PNCs; the inset image is the photograph of a blue PNC-based LED under operation. Reproduced with permission.⁵ Copyright © 2015, John Wiley and Sons. (c–e) Energy level diagrams of blue PeLEDs. (d) EQE–current density plots for NiO_x and TFB/PFI based blue devices. Reproduced with permission.⁵⁴ Copyright © 2018, John Wiley and Sons.



prevent the interface emission quenching caused by PEDOT:PSS. Similarly, severe emission quenching was found at the $\text{NiO}_2/\text{CsPb}(\text{Br}_{1-x}\text{Cl}_x)_3$ interface (Fig. 4d), which can be suppressed by using TFB/PFI (perfluorinated ionomer) (Fig. 4e) instead of NiO_2 .⁵⁴ Such a modification can improve the EQE from 0.03% to 0.5% for deep blue PeLEDs (Fig. 4f).⁵⁴

To realize efficient blue PNC-LEDs, the ligand density and surface defects of PNCs also need to be optimized for better charge injection. Pan and co-workers performed surface engineering for ligand density optimization to achieve better charge injection.⁵² This was enabled by treating the synthesized PNCs with oleylamine and oleic acid using DDABC (di-dodecyl dimethyl ammonium bromide chloride) *via* the ligand exchange strategy. Upon this, they achieved a sky-blue PeLED (the device structure is the same as that in Fig. 4c) with a turn-on voltage of 3.0 V, an EQE of 1.9%, and a luminance of 35 cd m^{-2} . The relatively low turn-on voltage suggests that the charge injection is improved after ligand density optimization. However, the ligand exchange process might change the crystallinity and crystal arrangement of PNCs, leading to cracks and defects in their thin films and hence reduce the device performance.¹⁷ Li *et al.* reported a cross-linking method to replace a long chain ligand by a short chain ligand using TMA (trimethylaluminum) treatment.¹⁷ This method will not affect the PNC arrangement and can lead to stable PNCs. A blue PeLED was thus fabricated using a device structure of ITO/ZnO/CsPb(Br_{1-x}Cl_x)₃/TFB/MoO₃/Au, showing an EQE of 0.007% and a luminance of 8.7 cd m^{-2} .¹⁷ In addition, they observed an electrical-induced phase segregation effect in

their blue device, the EL peak of which shifts towards green emissive Br-rich domains with increasing time intervals at a constant applied voltage. Phase segregation is a common phenomenon in mixed halide perovskites, which takes place at grain boundaries where a higher concentration of positive space charge possibly provides the initial driving force.^{55,56} It is a serious stability issue that should be addressed in blue PeLEDs as most of the reported blue emissive perovskites are achieved based on the mixed-Br/Cl strategy. Possible strategies that can improve the colour stability will be discussed later.

In addition to the control of the organic ligands and surface defects, manganese (Mn) doping in PNCs can increase their PLQEs and thus improve the EQE of blue PeLEDs.⁵⁷ Despite the fact that Mn-doped PNCs are reported with an additional emission peak from Mn^{2+} ,⁵⁸ a small amount of doping will not influence the emission colour; instead, it can increase the PLQE from 9% to 28% (Fig. 5a).⁵⁷ They finally fabricated a bright PeLED with a luminance of 400 cd m^{-2} and an EQE of 2.12% (Fig. 5b and c). It is also worthwhile to note that Mn-doped devices show narrower EL spectra compared to the control device (Fig. 5d).

3. Challenges and future directions

3.1 Colour stability

Blue emissive perovskites are prepared from mixed halide (Br/Cl) ions or low-dimensional NCs with the quantum confinement effect. From the material stability point of view, the

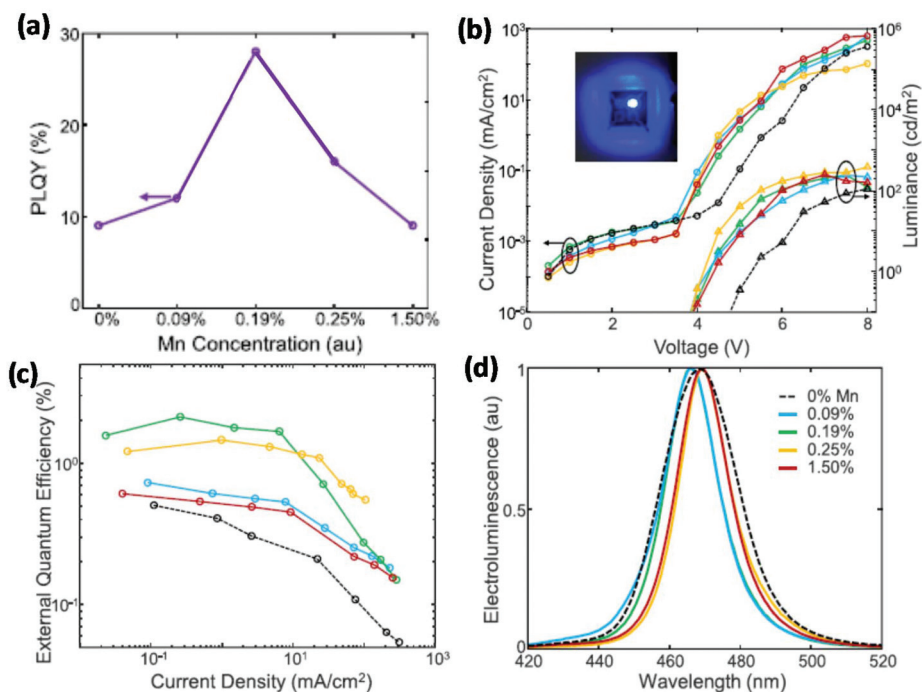


Fig. 5 Blue PeLEDs based on Mn-doped PNCs with various doping concentrations. (a) PLQEs of Mn-doped PNCs. (b) J - V - L characteristics; the inset image is a photograph of a blue LED under operation. (c) EQE-current density plots. (d) EL spectra. Reproduced with permission.⁵⁷ Copyright © 2018 Elsevier Inc.



phase segregation and moisture stability of mixed halide perovskites need to be investigated in depth. These materials show reversible electrical-field-induced phase segregation in blue PeLEDs.¹⁷ As shown in Fig. 6a and b, the blue EL peak shifts towards green (from Cl-rich to Br-rich domains) wavelengths under operation of the device at a constant bias voltage of 5.5 V; after resting for 20 min, the blue PeLED recovered blue emission. Metal stearate (MSt), which works as a capping agent that binds to the surface of PNCs, was successfully reported to prevent potential damage during the PNC purification process.⁵⁹ It was also reported that anion exchange between CsPbBr₃-MSt and CsPbI₃-MSt nanocomposites was greatly suppressed. This suggests that the efficient coating of MSt can effectively isolate the contact between CsPbBr₃ and CsPbI₃ nanocrystals, reflecting notable compositional stability.⁵⁹ As a result, stable CsPb(I_{1-x}Br_x)₃ PNCs were realized *via* impregnating them in MSt, which show a good colour stability under ambient conditions (Fig. 6c). Moreover, solvents with high dielectric constants can damage PNCs and create ionic defects during the purification of PNCs.^{14,60} A careful selection of the solvents can lead to stable PNCs. For instance, Chiba *et al.* used ethyl acetate (AcOEt) to wash long chain ligands such as alkyl-based oleylammonium iodide (OAM-I) and aryl-based aniline hydroiodide (An-HI) of

CsPb(I_{1-x}Br_x)₃ PNCs, upon which PeLEDs were fabricated with stable EL spectra peaking at 650 nm under various current densities (Fig. 6d).¹⁴ Similar methods can be used to synthesize PNCs for blue PeLEDs with good colour stability.

In addition, Braly *et al.* replaced the MA cation by FACs cations and demonstrated stable emission from mixed bromide-iodide perovskites.⁵⁵ It was reported that FACs-based mixed halide perovskites will ultimately show similar phase segregating behaviour but they can be more stable than MA-based perovskites. A combination of larger morphological grains and the presence of size mismatch on the perovskite monovalent cation site in FACs-based perovskites stabilize these materials.⁵⁵ Similar approaches can be developed to suppress the electrical-field-induced phase segregation in mixed-Br/Cl perovskites for colour-stable blue emission. One successful example was reported for triple cation based blue PeLEDs, which show stable EL spectra at different bias voltages (Fig. 1d).²² Furthermore, as mentioned above, the careful selection of the organic spacer materials for quasi-2D perovskites can also lead to colour-stable blue PeLEDs (Fig. 2d-f).^{38,39} Another plausible way is to decrease the dimensionality of pure bromide perovskites that feature the quantum confinement effect to realize blue emission. Such an approach has been successfully applied in pure iodide perovskites that show

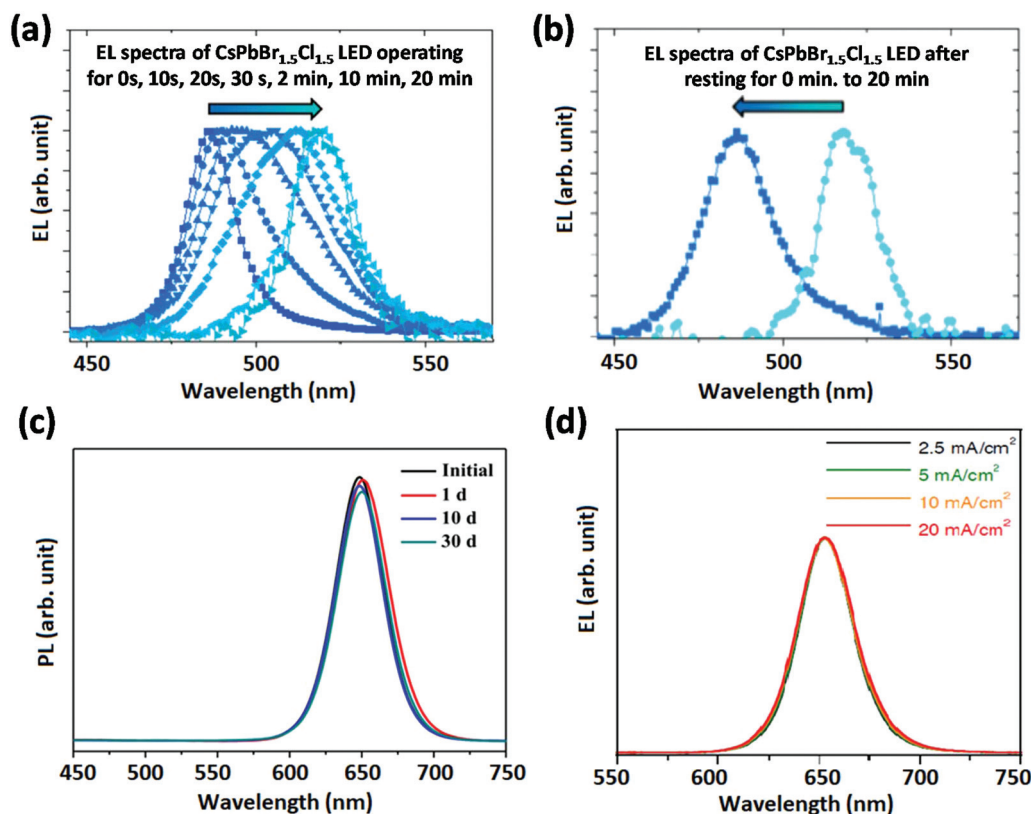


Fig. 6 Colour stability of blue PeLEDs. EL spectra recorded (a) after different time intervals and (b) after resting for 0/20 min. Reproduced with permission.¹⁷ Copyright © 2016, John Wiley and Sons. (c) PL spectra of CsPb(I_{1-x}Br_x)₃ PNCs with a metal stearate solid matrix. Reproduced with permission.⁵⁹ Copyright © 2018, American Chemical Society. (d) EL spectra under various current densities. Reproduced with permission.¹⁴ Copyright © 2018, Springer Nature.



colour-stable red emission (667 nm) tuned from the near-infrared region (790 nm) by decreasing the crystal size of perovskite films.⁶¹

3.2 Photoluminescence quantum efficiency

The improvement of the PLQEs of blue emissive perovskite films is necessary for realizing highly efficient blue PeLEDs as the EQE of a PeLED is decisively determined by the PLQE of its emissive layer and light out-coupling efficiency.^{16,24} Wider bandgap materials have more sub-bandgap states and a high defect density as compared to smaller bandgap materials,⁶² raising difficulties for obtaining blue emissive perovskites with high PLQEs. In addition, self-quenching of PNCs during deposition into thin films from colloidal solutions, which is mainly due to the formation of surface defects, should be suppressed.⁶³ A possible way to improve the PLQEs of blue emissive PNCs is the doping of metal ions such as Mn^{2+} , which can passivate non-radiative recombination centers.^{64,65} A successful example was reported by Hou *et al.*, where the PLQE of blue $CsPb(Br_{1-x}Cl_x)_3$ PNCs was increased from 9% to 28% (Fig. 5a).⁵⁷ Metal halides (*e.g.*, ZnX_2) can also help to suppress sub-bandgap states.⁴⁵ Similar approaches can be used for improving the PLQEs of blue emissive perovskites. For 2D blue perovskites, the suppression of exciton-phonon coupling can largely increase the PLQE up to 79%.⁶⁶ Passivation of defects at grain boundaries and film surfaces can also increase the PLQEs of perovskite films.⁶⁷ Morphology control by additives such as PEO (poly(ethylene oxide)) and crown ethers can largely promote the PLQEs of green emissive perovskites.⁶⁸⁻⁷⁰ These approaches could also be valid for blue perovskites.

3.3 Device structure

Blue emissive perovskites such as $CsPb(Br_{1-x}Cl_x)_3$ have a deep valence band maximum (≈ 6.1 eV) (Fig. 7),⁵⁴ which can lead to a large charge injection barrier for blue PeLEDs. Wide bandgap hole-transport layers (HTLs) with matching energy levels are needed to facilitate hole injection in blue PeLEDs. In addition, emission quenching at the interface between the HTL/ETL and the perovskite layer should be minimized. Gangishetty *et al.* found that the TFB/PFI layer shows less emission quenching than the NiO layer for $CsPb(Br_{1-x}Cl_x)_3$.⁵⁴ Hence, the EQE was improved from 0.03% to 0.5% for blue PeLEDs (Fig. 4f). Similarly, Zhang *et al.* found that the insertion of a thin layer of PVP (poly(vinyl pyrrolidone)) between ZnO and the green emissive perovskite layer can realize an enhancement in PLQE, leading to an enhanced EQE from 0.09% to 2.41% for $CsPbBr_3$ based PeLEDs.⁷¹ Considering that blue emissive perovskites have wide bandgaps, similar approaches should be taken into consideration to suppress emission quenching by the HTL/ETLs with lower bandgaps or even sub-bandgap states. Some possible materials with energy levels are shown in Fig. 7. In addition, p-type-doped HTLs and n-type-doped ETLs can be used to reduce the charge injection barrier at the interface between the HTL/ETL and perovskite and tune the charge transport properties,⁷² and thus improve hole/electron injection and transport balance to realize high EQEs. For example, CBP (4,4-bis(*N*-carbazolyl)-1,1-biphenyl) or PVK can be doped with MoO_x to improve hole injection in blue PeLEDs. In addition, a dipole functional interlayer such as PEI (polyethylenimine), PEIE, PFI *etc.*, can be used to tune the work function of HTL/ETLs to get better charge injection.⁷³⁻⁷⁵

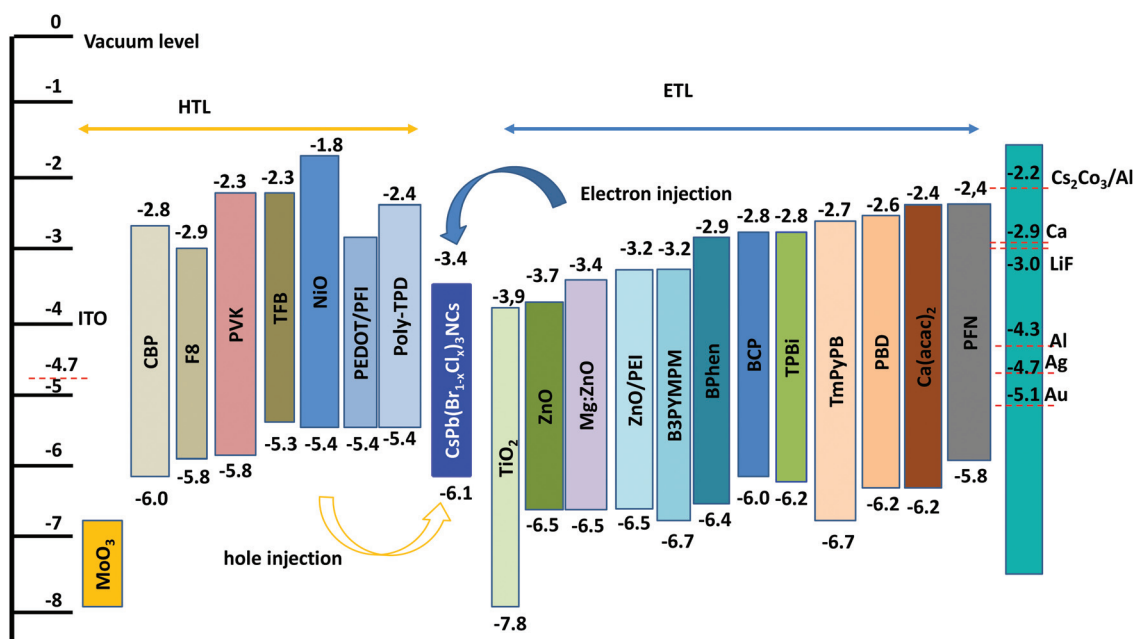


Fig. 7 Energy levels of various hole (HTL) and electron-transport layer (ETL) materials with blue emissive PNCs.



4 Conclusion and outlook

We have summarized the current scenario of blue PeLEDs and discussed the challenges in terms of colour instability, poor PLQE, and device structure. Plausible strategies are provided to stabilize the colour stability and improve the PLQE of blue emissive perovskites. Lowering the dimensionality of pure bromide perovskites can be a very plausible way to realize efficient and colour-stable blue emissive perovskites. This can be performed in combination of defect passivation of perovskites using additives. In addition, emission quenching at the interfaces between the HTL/ETL and perovskite layer should be minimized. Wide bandgap HTL/ETLs with negligible sub-bandgap states are thus needed. Alternatively, a thin interlayer of insulators such as PFI and PVP can be used to avoid emission quenching by the HTL/ETLs. In addition to efficiency, the device operational lifetime is also a major concern for blue PeLEDs. It will be desirable for future displays and lighting applications to design PeLEDs with a high efficiency and excellent stability. We believe that this can be realized by elaborate design of new materials and device architectures.

Conflicts of interest

There are no conflicts to declare.

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