

Cite this: *RSC Adv.*, 2017, 7, 34875

Nanomaterials for photocatalytic hydrogen production: from theoretical perspectives

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To overcome the increasing demand of energy worldwide and global warming due to CO₂ emissions from the use of traditional fuel sources, renewable and clean energy sources are in high demand. Solar energy is one of the important renewable energy sources since it can be converted into hydrogen fuel *via* water splitting. To produce highly efficient and low-cost H₂ from the reduction of water and CO₂, photocatalytic materials have been developed. Presently, the efficiency of H₂ production using transition metal-oxide and non-metal oxide photocatalysts *via* water splitting is quite low. The main issues reported are low light absorption and poor charge separation. The reasons for these two issues are the large band gaps and band edge positions of the present photocatalysts used for H₂ production. To produce H₂ to a larger extent from the decomposition of water, the present photocatalysts have been modified by co-catalysts or dopants using different techniques, including the reduction of the band gap and adjustment of the morphology, band edge positions, crystallinity, surface structure, *etc.*, such that these photocatalysts can absorb sufficient light in the visible-light region. This type of modified nanostructured photocatalysts (both oxide and non-oxide) can enhance the efficiency of H₂ production *via* absorbing sufficient light in the visible-light region of the solar spectrum and improving the charge separation by suppressing charge recombination. In this regard, we reviewed both UV- and visible-light active nanostructured photocatalysts and modified photocatalysts reported in theoretical studies.

Received 24th March 2017
Accepted 6th June 2017

DOI: 10.1039/c7ra03435k

rsc.li/rsc-advances

1. Introduction

The use of traditional energy sources, such as fossil and mineral fuels and nuclear and hydroelectric sources, causes global warming due to the emission of CO₂, methane, CO, *etc.* that can cause serious disasters in various places worldwide. Moreover, there is a continuous depletion of traditional energy sources. To solve these issues, alternative energy sources, such as renewable energy sources, have been developed by researchers worldwide to meet the world's energy needs as well as to decrease pollution.

Various types of renewable energy sources are under development or have been developed. Among these, hydrogen can be a promising potential candidate and its production from renewable energy sources is CO₂-free. In this regard, hydrogen can be considered as a clean energy carrier that can address the current energy and environmental issues.^{1,2}

Hydrogen can be produced from primary energy sources including coal, natural gas, waste, biomass, solar, wind, hydropower, nuclear power, and geothermal power. For example, commercial hydrogen can be produced *via* steam reforming of methane^{3,4} and also from coal using gasification

technology; however, both these methods result in CO₂ emissions. Hydrogen can be stored in its gaseous, liquid, or metal hydride form and can be distributed over large distances *via* pipelines or tankers. Numerous studies have been carried out to discuss the key role of hydrogen in sustainable development.^{5–10}

The abundant supply of water and sunlight offers us an affordable alternative source to produce hydrogen. Therefore, photocatalytic water splitting is one of the potential techniques for clean solar hydrogen production and has been utilized in small- to large-scale hydrogen generators.¹¹ In 1972, Honda and Fujishima¹² first investigated water splitting using a single TiO₂ crystal as a photoanode and Pt as a cathode. Subsequently, numerous studies on water splitting have been conducted^{13–19} with the exploration of more than 100 different new catalysts including multi-component oxides, sulfides, nitrides, and carbides. All of them are concentrated on the various factors affecting the conversion efficiency such as how much sunlight is absorbed, the exciton generation, the separation/recombination of holes and electrons received by dissociation of excitons, and the charge collection by the respective electrodes to produce hydrogen and oxygen. Some other practical factors are also important for photocatalytic hydrogen production such as stability and resistance under visible light, reduced cost, and non-toxicity of utilized materials.^{20–23} In this regard, we speculated that photocatalytic hydrogen could be a commercial fuel

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in future; however, significant research efforts should be devoted in this direction, addressing the abovementioned key issues for its industrial production.

Nanomaterials such as CdS, SiC, CuInSe₂, and TiO₂ can be utilized to produce cheap and clean photocatalytic hydrogen^{24–27} due to the fact that these nanomaterials show better photocatalytic properties as compared to their bulk counterparts. Currently, many other nanomaterials such as Nb₂O₅,²⁸ Ta₂O₅,²⁹ α -Fe₂O₃,^{30,31} ZnO,^{32,33} TaON,³⁴ BiVO₄,^{35,36} and WO₃ have been explored.³⁷ In most of the photocatalysts, band gap limitation is a key issue that results in low H₂ production.³⁸ To solve this issue, some photocatalysts have been modified *via* noble metal doping, ion doping, sensitization, and metal-ion implantation. In noble metal doping, Pt has been found to be the best noble metal; however, it is too expensive. Thus, other efficient and cheaper metals such as Ni, Cu, Ag, Ru, Pd, and Ir have been explored.^{39–47} Ion doping involves transition metal and rare-earth metal ions, anions of nitrogen and sulphur, *etc.* Sensitization means dye-sensitization and coupling of semiconductors. Among the abovementioned modification techniques, metal-ion implantation and dye-sensitization have been found to be the most effective photoanode surface modification techniques. Moreover, researchers are interested in exploring co-catalysts with photocatalyst nanomaterials that can be used for photocatalytic hydrogen production.

The theoretical and computational description of many body systems is still one of the biggest challenges in solid state science although significant progress has been made in this regard. The majority of theoretical and computational studies on photocatalytic nanomaterials have used density functional theory (DFT) because DFT predicts the ground-state properties in terms of electronic density. The band gap of the materials is often significantly underestimated by DFT,⁴⁸ computed using Kohn–Sham method by introducing an exchange correlation (XC) functional such as local density approximation (LDA)⁴⁹ or the generalized gradient approximation (GGA).⁵⁰ The band gap of the materials can be improved by applying several approaches. For example, a hybrid functional can improve the accuracy by combining the approximate DFT functionals (LDA or GGA) with the exact Hartree–Fock (HF) exchange energy.⁵¹ Alternatively, the DFT + U method with the Hubbard band U term can provide more accurate results.⁵² The accuracy of these two approaches depends on the selection of several empirical parameters. Moreover, the parameter-free GW method yields an accurate description of the electronic structure of the materials, but at a huge computational expense.⁵³ Therefore, the use of GW method is not more reliable for complex structures.

Herein, the study was concentrated on modeling a nanostructured photocatalyst and modified nanophotocatalysts at the atomic level to provide information on the key atomic level structures, processes, and parameters that control the behavior of photocatalytic materials during their applications in photocatalytic H₂ production. In particular, substantial contributions have been made by modeling methods in the study of the electronic states, defects, and surface properties of nanophotocatalysts for H₂ production.

2. Theoretical achievements for the development of nanophotocatalysts and modified photocatalysts for hydrogen production *via* water splitting

Many nanophotocatalysts have been developed to split water into H₂ and O₂ under UV and visible-light illumination. The industrial production of clean and recyclable H₂ *via* the direct splitting of water using a particulate photocatalyst could be the best method.^{54–57} However, there is a lack of suitable materials with appropriate band gap positions for overall water splitting, and stability is required for practical applications. The reason is that most metal oxides have optical band gaps that lie outside the visible-light range ($E_g > 3$ eV); thus, even if they are catalytically active for a given reaction, they cannot make efficient use of the solar spectrum. For example, TiO₂ has a wide band gap ($E_g = 3.2$ eV), which is limited to a 1% solar-to-hydrogen (STH) conversion efficiency for the water splitting process, whereas for a smaller band gap material such as Fe₂O₃ ($E_g = 2.2$ eV), the theoretical efficiency increased to 15%.⁵⁸ The STH of 10% is required for the solar cell-powered electrolysis of water. Therefore, the development of highly active photocatalysts should be explored for large-scale photocatalytic H₂ production *via* water splitting.^{59–62} There are two types of nanostructured photocatalyst materials for H₂ production *via* water splitting.

(i) Metal-oxide photocatalysts and modified metal-oxide photocatalysts (with hybrid systems)

Although there have been many experimental investigations reported on metal oxide nanophotocatalysts for H₂ production *via* water splitting under UV- and visible-light irradiation, very few theoretical investigations based on the DFT methods have been reported. Among these, most theoretical investigations have been reported on nanostructured TiO₂ and modified TiO₂ photocatalysts. For example, Kaur *et al.*⁶³ investigated amorphous TiO₂ as a photocatalyst for H₂ production using the DFT methods. The authors claimed *via* the analysis of the electronic properties that amorphous TiO₂ may act as a cheaper, more abundant, but somewhat less efficient photocatalyst as compared to crystalline TiO₂. To improve the photocatalytic activity of the TiO₂ nanophotocatalyst, the doping of TiO₂ was carried out using different techniques *via* narrowing the band gap of TiO₂ to absorb more visible-light in the solar spectrum as well the dependence of charge separation and recombination on the distance between the catalytic core and semiconductor surface. For example, Reynal *et al.*⁶⁴ reported that the photoinduced reduction of H₂ on a Co electrocatalyst immobilized on TiO₂ was 10⁴ times faster than the reverse charge recombination. The authors claimed that both processes show an exponential dependence on the distance between the semiconductor and the catalytic core, as shown in Fig. 1(a)–(d).⁶⁴ As shown in Fig. 1, the authors computed the charge separation and recombination processes when a semiconductor (TiO₂) was functionalized with three related cobalt electrocatalysts, whose



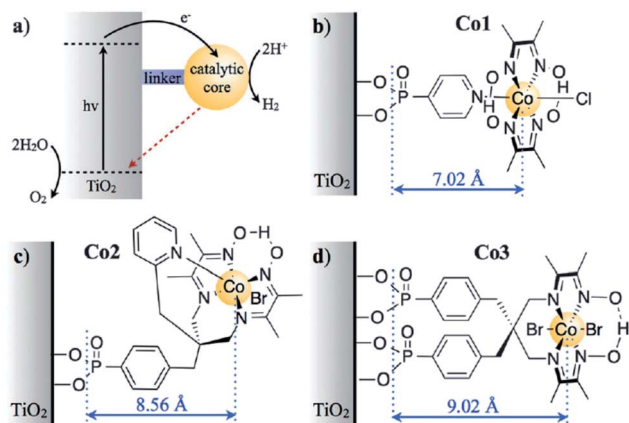


Fig. 1 (a) The electron transfer processes in TiO_2 functionalised with a molecular catalyst for H_2 production after UV-light excitation. The solid black and dashed red arrows indicate the charge separation and recombination, respectively. The molecular structures of the catalysts for H_2 reduction are shown in (b) for Co1, (c) for Co2, and (d) for Co3 (the charges have been omitted for clarity). The blue arrows indicate the distance between the anchoring groups and the catalyst metal centre (r , Å), as determined by the energy minimised DFT calculations. Reprinted with permission of ref. 64. Copyright 2014 Royal Society of Chemistry.

molecular structure varied the physical separation between the catalytic core and the semiconductor surface. In this hybrid system, the semiconductor acted as a light harvester and H_2 evolution was driven by the anchored molecular catalyst. Some theoretical investigations have shown that BaTiO_3 (ref. 65 and 66) and PbTiO_3 (ref. 67–69) nanowires exhibit ferroelectricity and $\text{Bi}_2\text{Ti}_2\text{O}_7$,⁷⁰ SrTiO_3 ,⁷¹ and CaTiO_3 (ref. 72) nanowires exhibit photocatalytic activity towards water splitting. Bandura *et al.*⁷³ demonstrated the photocatalytic activity of SrTiO_3 nanowires and concluded that *via* comparison of the band edge positions and Fermi levels relative to the levels of the reduction H^+/H_2 and oxidation $\text{O}_2/\text{H}_2\text{O}$ potentials as well as analysis of the band gap widths, the probable SrTiO_3 nanowire configurations and SrTiO_3 nanotube⁷⁴ configurations can be suggested for application in the photocatalytic splitting of water molecules under solar irradiation in the visible-light range.

In the case of TiO_2 , the phase of the photocatalyst is a key factor for its photocatalytic activity. For example, although the rutile phase exhibits a narrower band gap (approx. 3.0 eV) than the anatase phase (approx. 3.2 eV),^{75–77} anatase is generally considered to exhibit superior photocatalytic activity due to its larger surface area and thus higher activity.⁷⁸ The third metastable phase is brookite that has also been reported to exhibit photocatalytic activity,^{79,80} and mixtures of anatase, rutile, and brookite have been reported to exhibit significant photocatalytic activity.^{81,82} However, this phase is of lesser interest due to the complexity of its synthesis. The doping of TiO_2 can provide several outcomes: formation of new valence states in TiO_2 ,^{83–86} creation of charge carrier trapping sites,^{87–90} band gap reduction,^{91–93} control of phase transformation behavior,^{94–97} and surface enhancement.^{96,98} The effects of dopants on the anatase to rutile phase transformation have been comprehensively

reviewed elsewhere.⁹⁹ The phase transformation mainly occurs due to the atomic rearrangement involved in the transformation, which is often a result of an increase in the density of the anion vacancies. This may occur when cationic dopants of low valence substitute Ti in the anatase lattice.^{79,99,100} Hanaor *et al.*¹⁰⁰ performed DFT calculations for the relative stability of anatase and rutile polymorphs of TiO_2 . The rutile phase exhibited a stability of 3 meV as compared to the anatase phase in pristine TiO_2 . The authors also demonstrated the doping of TiO_2 with Si, Al, Fe, and F atoms and found that the anatase phase stabilized the rutile phase in the dopant order $\text{F} > \text{Si} > \text{Fe} > \text{Al}$ obtained *via* comparison of ΔF (a- TiO_2 :F) with ΔF (r- TiO_2 :F), which provided their formation energies as 84.78, 32.43, 18.46, and 17.68 meV, respectively. As the authors considered only fluorine anion dopant in their study, the inhibition of rutile formation *via* F doping was supported by previously reported experimental results.^{101,102} A schematic of the interstitial and substitutional dopant positions in anatase and rutile TiO_2 is shown in Fig. 2(a) and (b).¹⁰⁰

In general, efforts to modify TiO_2 to enhance the visible-light absorption have been focused on the substitution of metal cations and/or non-metal anions.^{103–111} In some cases, the choice of DFT methodology can have an influence on the reduction of the band gap to the visible-light region of the solar spectrum.¹⁰⁹ Co-doping with compensating cation–anion pairs from DFT simulations has been reported,^{112,113} along with a recent experimental study of Mo–C co-doped TiO_2 .¹¹⁴ Key parameters such as stability, solubility, and reproducibility are also important along with the band gap reduction as well as the method of doping and/or the DFT methodology to enhance the photocatalytic activity of TiO_2 due to minimum charge recombination. Moreover, nanostructure engineering of TiO_2 can reduce the charge recombination.^{115–119} In this regard, heterostructure techniques of TiO_2 with other oxides can enhance the photocatalytic activity to tune the structure to increase the absorption of visible-light.^{120–133} For example, heterostructures such as $\text{Bi}_4\text{Ti}_3\text{O}_{12}$ – TiO_2 ,¹²⁴ MgO – TiO_2 ,¹³⁴ and Ga_2O_3 – TiO_2 (ref. 134) can reduce the band gap, leading to visible-light photoactivity, efficient charge separation, and improved photocatalytic activity as compared to pure TiO_2 .

In the case of the $\text{Bi}_4\text{Ti}_3\text{O}_{12}$ – TiO_2 heterostructure, Bi 6s mainly contributes to the valence band, whereas the Bi 6p state contributes to the conduction band, which result in the reduction of the band gap of $\text{Bi}_4\text{Ti}_3\text{O}_{12}$ (2.5 eV) as compared to that of TiO_2 (3.2 eV). This leads to large absorption of visible-

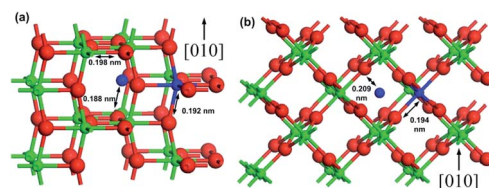


Fig. 2 A schematic of the interstitial and substitutional dopant positions in (a) anatase and (b) rutile TiO_2 with the arrows showing the cation-dopant to oxygen distances. Reprinted with permission of ref. 100. Copyright 2012 Springer-Verlag Berlin Heidelberg.



light. A schematic showing the energy band structure and electron-hole pair separation in the $\text{Bi}_4\text{Ti}_3\text{O}_{12}$ - TiO_2 heterostructure is shown in Fig. 3.¹²⁴ Similarly, in the case of MgO - TiO_2 and Ga_2O_3 - TiO_2 heterostructures, the author claimed the presence of two states in the band gap, as shown in Fig. 4.¹³⁴ The first state comes from an occupied Ti^{3+} state at 0.9 eV and 1.0 eV above the valence band edge in the MgO - TiO_2 and Ga_2O_3 - TiO_2 heterostructures, respectively. The second state comes from an unoccupied O 2p state (oxygen hole polaron), which lies 1.8 eV and 1.5 eV above the valence band of the MgO - TiO_2 and Ga_2O_3 - TiO_2 heterostructures, respectively. These results show that the excitonic electrons and holes are localized on the TiO_2 and the metal oxide nanostructures, which lead to a reduction in charge recombination and an enhancement in the photocatalytic activity.

In addition, TiO_2 nanoparticles can attach materials such as zeolites,¹³⁵ silica,¹³⁶ activated carbon^{137–142} or carbon nanotubes.^{143–149} Recently, graphene has attracted significant interest^{150–154} due to its large surface area and potentially higher photocatalytic efficiency.¹⁵⁵ However, very few reports are available for theoretical investigations on the interaction of nanoparticles and graphene sheets, and the mechanism of charge transfer from graphene to the nanoparticles. The decorated surface of graphene with both TiO_2 nanofibers¹⁵⁵ and Fe-doped TiO_2 nanofibers¹⁵⁶ in supercritical carbon dioxide exhibits high photocatalytic activity. Das *et al.*¹⁵⁷ deposited different semiconductor nanoparticles, such as TiO_2 and ZnO , and some magnetic nanoparticles, such as Fe_3O_4 and Ni, on graphene. They calculated their electronic structure using DFT methods and found that charge transfer occurred between graphene and the deposited nanoparticles.

Among all the metal oxides, WO_3 has attracted significant interest due to its photosensitivity,^{158–160} good electron transport properties,¹⁶¹ and stability against photocorrosion.^{162,163} Moreover, its smaller band gap (approx. 2.8 eV) than other oxides such as TiO_2 makes it suitable for the absorption of visible-light. However, the gap of WO_3 is still too large for sufficient visible-light absorption. Experimental results have shown that the conduction band minimum (CBM) of bulk WO_3 is about 0.4 eV below the hydrogen redox potential.^{164,165} For the WO_3 surface, the CBM lies 0.31 eV above the hydrogen redox potential, which is still low for hydrogen production.¹⁶⁶

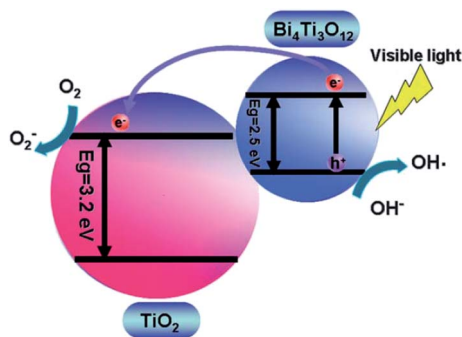


Fig. 3 A schematic showing the energy band structure and electron-hole pair separation in the $\text{Bi}_4\text{Ti}_3\text{O}_{12}/\text{TiO}_2$ heterostructure. Reprinted with permission of ref. 124. Copyright 2011 Royal Society of Chemistry.

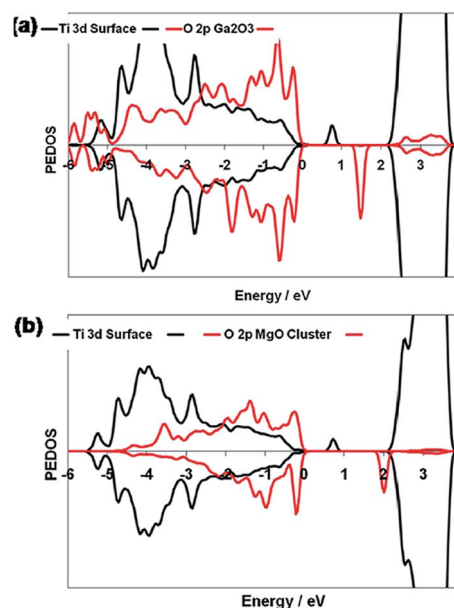


Fig. 4 Ti 3d and O 2p PEDOS of the $(\text{Ga}_2\text{O}_3)_2$ and $(\text{MgO})_4$ clusters for the relaxed triplet excited states of (a) Ga_2O_3 -modified TiO_2 and (b) MgO -modified TiO_2 . The zero of energy is the top of the valence band. Reprinted with permission of ref. 134. Copyright 2012 American Chemical Society.

Doping and co-doping are good ways to tailor the electronic structure and photocatalytic properties of semiconducting oxides. For example, Mo-doped WO_3 nanowires showed an enhancement in visible-light photoactivity, as the band gap of the $\text{Mo}_x\text{W}_{1-x}\text{O}_3$ solid solutions was narrowed by 0.48 eV upon increasing the Mo content from 0 to 0.75.^{167,168} The effect of several transition metals on the photocatalytic activity under UV irradiation was also studied, suggesting maximum H_2 production *via* Ni doping with WO_3 .¹⁶⁹ Doping with other metals (Ti, Zn, Dy, Te, Ta, V, Cu, Ag, and Ce) has also been reported in the literature,^{170–178} with some claims of improved photocatalytic activity. For example, Wang *et al.*¹⁷⁹ studied Mo, Cr, Ti, Zr, and Hf-doped WO_3 and found that in Hf-doped WO_3 , the oxygen vacancies have a negative formation energy, which leads to a shift in both the VB and CB to higher energies and a reduction of the band gap, with potential benefits for photocatalytic H_2 production, as shown in Fig. 5.¹⁷⁹

In previous studies, it has been reported that perovskite-type alkali tantalates, ATiO_3 ($A = \text{Li, Na, and K}$) show significant photocatalytic activity under UV irradiation. However, these are less active under visible-light irradiation.^{180,181} Transition metal ions were found to be attractive dopants for the effective modification of photocatalysts, which can play multiple roles ranging from reducing the band gap to efficient electron transfer due to their multiple oxidation states.^{182,183} Electron transitions from the valence band to the dopant level or from the dopant level to the conduction band can effectively red shift the band edge absorption threshold.

For example, Liu *et al.*¹⁸⁴ demonstrated M-doped NaTaO_3 ($M = \text{V, Cr, Mn, Fe, Ni, Cu, and Zn}$) nanoparticles using both experimental and DFT techniques. The authors found that the



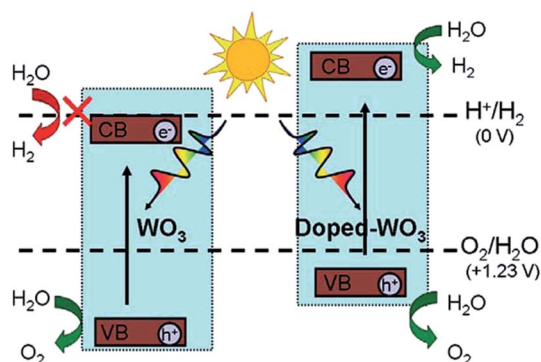


Fig. 5 A schematic of the superior photocatalytic activity of metal-doped WO_3 for H_2 production as compared to that of WO_3 . Reprinted with permission of ref. 179. Copyright 2012 American Chemical Society.

substitution of Ta^{3+} by metal ions results in the formation of an intermediate band, which is due to the metal 3d state. The supercell of M-doped NaTaO_3 , total and projected DOS of pure and V-doped NaTaO_3 , and intermediate band formation by the

d-orbitals of the doping metal ions are shown in Fig. 6(a), (b) and (c), respectively.¹⁸⁴

Moreover, a ZnO /graphene hybrid nanostructure was reported to have excellent potential for application in electronic devices.^{185–190} Most studies on the ZnO /graphene hybrid structures have been focused on their structural morphology and electronic properties.^{191,192} Recently, Xiang *et al.*¹⁹³ first proposed a three-component composite, TiO_2 /graphene/ MoS_2 , containing 0.25 wt% graphene. The results showed that graphene acts as an electron reservoir and MoS_2 can act as a source of active adsorption sites to achieve a highly efficient synergetic H_2 evolution at the rate of $165.3 \mu\text{mol h}^{-1}$, which is 17 times that of $\text{TiO}_2/\text{MoS}_2$.¹⁹⁴ Wu *et al.*¹⁹⁵ studied ZnO -dotted porous ZnS cluster microspheres for highly efficient, Pt-free photocatalytic H_2 evolution.

(ii) Non-metal oxide photocatalysts and modified non-metal oxide photocatalysts (with hybrid systems)

The surface modification of oxide photocatalysts, particularly TiO_2 , via doping with N, C or S, or metal nanoparticles cannot provide significant photocatalytic activity in the modified

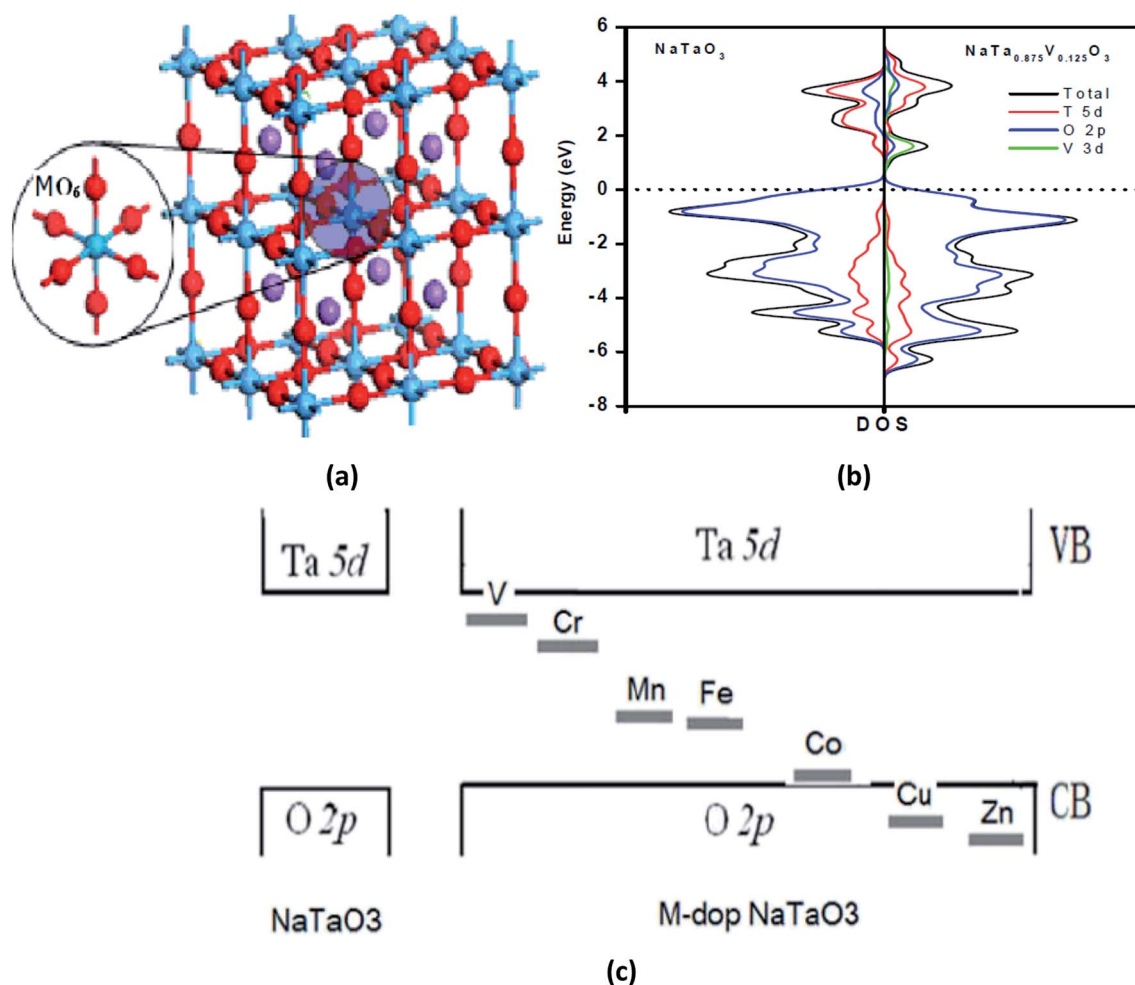


Fig. 6 (a) A supercell of M-doped NaTaO_3 $2 \times 2 \times 2$ (40-atoms), (b) the total density of states and projected density of states of the pure and V-doped NaTaO_3 , and (c) intermediate band formation by the d-orbital of the doped metal ions. Reprinted with permission of ref. 184. Copyright 2012 IACSIT Press, Singapore.



materials in the visible-light region of the solar spectrum.^{196,197} Therefore, researchers have been interested in exploring alternative materials for nanophotocatalysts and modified photocatalysts for photocatalysis and solar energy conversion. As a result, the large-scale production of stable visible-light active photocatalysts remains a challenge for industrial applications.¹⁹⁸ For example, polymeric carbon nitride has been found to be an efficient photocatalyst that produces H₂ from water under visible-light irradiation; however, in this case, a sacrificial donor is required.¹⁹⁹ Similarly, a class of metal-free photocatalysts, including elemental boron,²⁰⁰ sulfur,²⁰¹ and phosphorus,²⁰² as well as the binary carbon nitride²⁰³ and boron carbide, has emerged.²⁰⁴ Conjugated polymer nanostructures have emerged as alternate materials for applications in solar energy conversion. However, there is a lack of photocatalytic studies of these conjugated polymers. Semiconductor nanostructures modified with conducting polymers have been studied for photocatalytic applications. Two-dimensional materials such as graphene²⁰⁵ and layered hexagonal (h-BN) are promising photocatalysts.^{206–208} However, graphene (monolayer) exhibits a zero band gap, whereas h-BN exhibits a wide band gap (~5.5 eV). These materials (say ternary B–C–N compounds) can constitute the desired medium band gap semiconductors *via* adjustment of the band gap and absolute energy levels *via* chemical variations.^{209,210} Similarly, graphene oxide based on a carbon support may improve the charge separation and spontaneous redox processes.^{211,212} In recent years, graphene-based heterogenous photocatalytic

nanomaterials have attracted significant attention due to their unique sp² hybrid carbon networks, exhibiting ultra-fast electron mobility at room temperature, conductivity, large theoretical surface area, high work function, *etc.*^{213–215} Huang *et al.*²¹⁵ demonstrated a ternary structure of B–C–N nanosheets that has the functionality to catalyze H₂ and O₂ evolution from water as well as CO₂ reduction under visible-light illumination. The plane-wave DFT calculations of the electronic structure of h-BN are shown in Fig. 5(a) and (b).²¹⁵ As shown in Fig. 7(a) and (b), for a B₁₁C₁₂N₉ compound, the gap is significantly reduced from 4.56 eV to 2.00 eV. The partial DOS, as shown in Fig. 7(b), signifies that the VB and CB edges of B₁₁C₁₂N₉ mainly comprise C 2p orbitals, which is different from that of pure h-BN (Fig. 7(a)).

Metal chalcogenides have been found to be promising photocatalyst materials for photocatalytic H₂ production due to their appropriate band gap width and band edge position.²¹⁶ In particular, CdS has been found to be an efficient semiconductor H₂-production photocatalyst.²¹⁷ However, some key issues such as the quick recombination of photogenerated charge carriers and photocorrosion under visible-light irradiation still exist and prohibit the wide application of CdS to a large extent. To solve these issues, forming Zn_{1–x}Cd_xS solid solutions is a viable method because ZnS possesses the same coordination mode with CdS,^{218–220} and the band gap width and band edge position of the Zn_{1–x}Cd_xS solid solutions can be tuned by changing the molar ratio of ZnS and CdS.^{221,222} For example, Li *et al.*²²³ demonstrated the visible-light photocatalytic H₂ production activity of Zn_{1–x}Cd_xS solid solutions with different molar ratios of ZnS and CdS using both experimental and DFT methods. The authors found that for Zn/Cd with a molar ratio equal to 1 : 1, the Zn_{0.5}Cd_{0.5}S solid solution exhibits the highest H₂-production rate of 7.42 mmol h^{–1} g^{–1}, exceeding that of the pure CdS and ZnS samples by more than 24 and 54 times, respectively. The supercell models for bulk ZnS and CdS, the geometry of Zn_{0.5}Cd_{0.5}S, the band structures of the Zn_{1–x}Cd_xS systems for various Zn/Cd molar ratios and the conduction and valence band edge potentials of the samples (*x* = 0, 0.1, 0.3, 0.5, 0.7, 0.9, and 1.0) are shown in Fig. 8(a), (b) and (c), respectively.²²³ Fig. 8(d) shows that the band gap of the Zn_{1–x}Cd_xS samples gradually becomes narrower from 1.921 to 1.141 eV upon increasing the Cd content, which is well consistent with the experimental results.

Many studies have been conducted on the development of semiconducting material/graphene hybrid structures using either gas phase^{223–226} or liquid phase techniques.^{185,186,191,227} Hou *et al.*²²⁸ have developed a CdS QDs/graphene/ZnIn₂S₄ system that exhibits highly efficient H₂ production due to the high hydrothermal stability and efficient electron transfer. Graphene can also play a role as a co-catalyst and constitutes a synergistic effect with the other co-catalyst to improve the photocatalytic efficiency.^{229–231} Similarly, Zhu *et al.*²³⁰ demonstrated ZnS loaded with 0.25 wt% graphene and 2 atom% MoS₂ that yielded a high H₂ production rate of 2258 μmol h^{–1}, which is 2 times of that observed with ZnS only.

Consequently, there is a huge scarcity of theoretical investigations based on the DFT methods on nanophotocatalysts (both

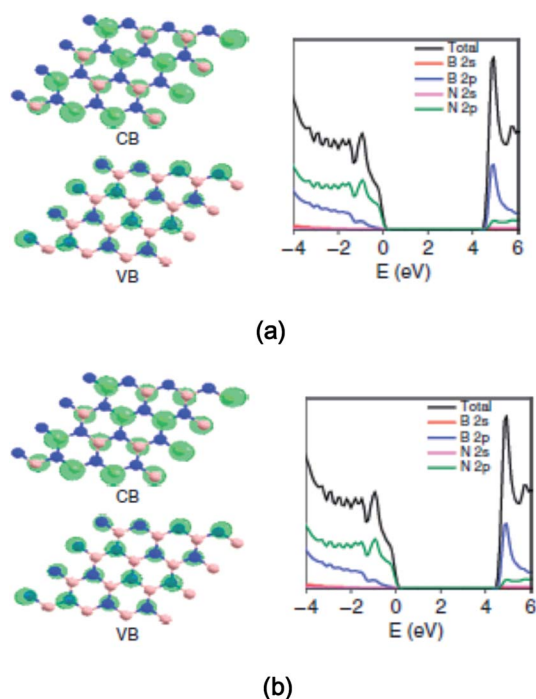


Fig. 7 The plane-wave DFT calculations of the electronic structure of h-BN and h-BCN. The optimized structure of B₁₆N₁₆ with the corresponding valence band (VB)/conduction band (CB) and the corresponding total and ion-decomposed electronic density of states (a and b). Reprinted with permission of ref. 215. Copyright 2015 Macmillan Publishers Limited.



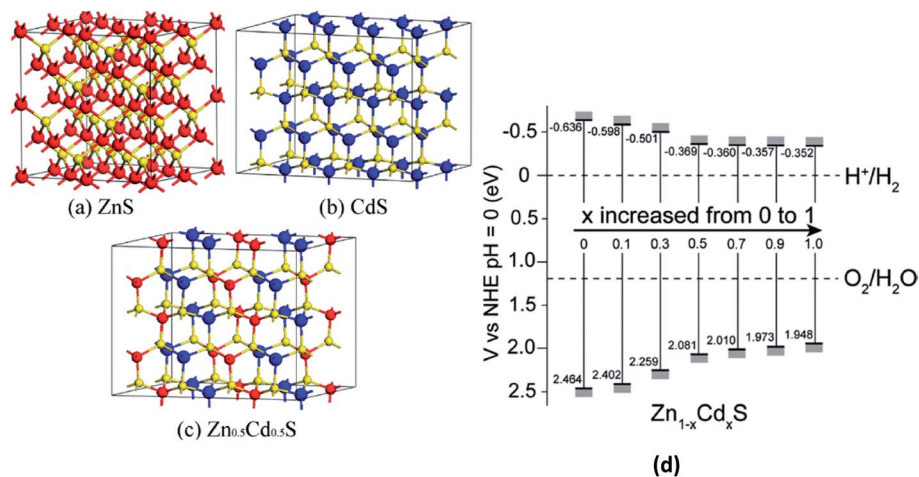


Fig. 8 The supercell models for bulk (a) ZnS (64 atoms) and (b) CdS (64 atoms). The red, yellow, and blue spheres represent the Zn, S, and Cd atoms, respectively. (c) The geometry model for Zn_{0.5}Cd_{0.5}S and (d) conduction and valence band edge potentials of the Zn_{1-x}Cd_xCd_{0.5}S samples ($x = 0, 0.1, 0.3, 0.5, 0.7, 0.9$, and 1.0). Reprinted with permission of ref. 223. Copyright 2013 American Chemical Society.

metal-oxide and non-metal oxide) and modified photocatalysts for photocatalytic H₂ production *via* water splitting. In the case of metal-oxide photocatalysts, few studies were reported on pure TiO₂ and modified TiO₂, which were focused on the electronic structures of modified TiO₂, signifying the gap reduction and efficient charge separation due to doping of TiO₂ to enhance the photocatalytic H₂ production. Moreover, very few studies have been reported for graphene and metal chalcogenides in the form of B–C–N ternary compounds and multicomponent heterostructures (metal oxide/graphene/metal chalcogenides) for efficient photocatalytic H₂ production due to the synergistic effect of graphene and MoS₂. There is a big opportunity for theoretical researchers to conduct DFT calculations on metal oxide and non-metal oxide photocatalysts and their modified forms for highly efficient photocatalytic H₂ production. Therefore, more DFT investigations should be devoted to either experimentally explore photocatalysts and modified photocatalysts or on the exploration of new photocatalytic materials for highly efficient H₂ production to a large extent.

3. Summary and perspectives

Herein, we reviewed nanostructured photocatalysts (both transition metal oxides and non-metal oxides) and modified photocatalysts for highly efficient H₂ production *via* water splitting under both UV and visible-light irradiation investigated *via* DFT techniques. However, the current lack of industrial applications of current semiconductor-based photocatalytic H₂ production is mainly due to two reasons: the low photocatalytic efficiency and lack of extensive studies for successful industrial applications of photocatalytic H₂ production experimentally as well as the lack of DFT studies to investigate the factors responsible for the enhancement of the efficiency of H₂ production. The current H₂ production yield is quite low, which is far from the targeted quantitative efficiency of 30% at 600 nm for practical applications.²³² The critical light conversation efficiency for photocatalytic H₂ production *via* water splitting is about 15%.²³³ In

this regard, the present research does not explore the highly efficient and cost-effective photocatalysts and modified photocatalysts for H₂ production *via* water splitting. The reason is that the present metal oxide and non-metal oxide photocatalysts have wide band gaps, which do not absorb sufficient sunlight in the visible-light region. To reduce the band gaps towards visible-light absorption and photogenerated charge separation in photocatalysis, the photocatalysts have been modified using different techniques in different forms. However, the modified photocatalysts do not enhance the efficiency of H₂ production from the decomposition of water under visible-light irradiation. Moreover, the fast charge recombination and fast back reactions are also drawbacks for the solar water splitting system.^{234–236} DFT investigations have reported the control of band gap and band edge positions *via* variations of the morphology, composition, crystallinity, and surface structure of photocatalysts for efficient H₂ or O₂ evolution from water splitting.

Consequently, more theoretical studies should be devoted towards the exploration of highly efficient, stable, eco-friendly and cost-effective photocatalysts, and modified photocatalysts *via* the continual addition of electron donors to overcome the key issues of current solar water splitting systems in the production of H₂ to a larger extent.

Acknowledgements

This work was supported by the Korea Center for Artificial Photosynthesis (KCAP, 2009-0093880, 2009-0093886), Basic Science Research Program (No. 2012-017247), BK Plus Program and A3 Foresight Program, all funded by the Ministry of Science, ICT and Future Planning through the National Research Foundation (NRF) Korea.

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