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## Enzyme/inorganic nanoparticle dual-loaded animal protein/plant protein composite nanospheres and their synergistic effect in cancer therapy†

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It is a viable strategy to develop a safer and tumor-specific method by considering the tumor microenvironment to optimize the curative effect and reduce the side effects in cancer treatment. In this study, glucose oxidase (GOx) and Fe<sub>3</sub>O<sub>4</sub> nanoparticles were successfully loaded inside regenerated silk fibroin/zein (RSF/zein) nanospheres to obtain dual-loaded Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres. The unique structure of the RSF/zein nanospheres reported in our previous work was favorable to loading sufficient amounts of GOx and Fe<sub>3</sub>O<sub>4</sub> nanoparticles in the nanospheres. For Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres, GOx depletes endogenous glucose *via* an enzyme-catalyzed bioreaction, simultaneously generating plenty of H<sub>2</sub>O<sub>2</sub> *in situ*. It was further catalyzed through a Fe<sub>3</sub>O<sub>4</sub>-mediated Fenton reaction to form highly toxic hydroxyl free radicals (•OH) in the acidic tumor microenvironment. These two successive reactions made up the combination of starvation therapy and chemodynamic therapy during cancer treatment. The catalytic activity of GOx loaded in the RSF/zein nanospheres is similar to that of the pristine enzyme. It was maintained for more than one month due to the protection of the RSF/zein nanospheres. The methylene blue degradation results confirmed the sequential reaction by GOx and Fe<sub>3</sub>O<sub>4</sub> from Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres. The *in vitro* experiments demonstrated that the Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres entered MCF-7 cells and generated •OH free radicals. Therefore, these Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres exhibited a considerable synergistic therapeutic effect. They showed more efficient suppression in cancer cell growth than either single-loaded GOx@RSF/zein or Fe<sub>3</sub>O<sub>4</sub>@RSF/zein nanospheres, achieving the design goal for the nanospheres. Therefore, the Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres cut off the nutrient supply due to the strong glucose dependence of tumor cells and generated highly toxic •OH free radicals in tumor cells, effectively enhancing the anticancer effect and minimizing side effects. Therefore, in future clinical applications, the Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres developed in this study have significant potential for combining starvation and chemodynamic therapy.

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## Introduction

In cancer treatment, it is critical to avoid damage to healthy tissues and organs caused by non-specific therapies.<sup>1</sup> Recently, several novel treatment methods, including photodynamic,<sup>2,3</sup> photothermal,<sup>4</sup> and microwave<sup>5</sup> therapies, have emerged to enhance the therapeutic effect and decrease the side effects. However, despite the high-efficiency treatment performance, these strategies could also damage normal tissues and induce

tumor metastasis.<sup>6</sup> Therefore, researchers have shifted their attention to exploring the tumor microenvironment to develop a safer and more tumor-specific method.<sup>7–10</sup> Significant differences exist between normal and tumor sites in cell metabolism and the physical environment. The tumor sites possess complex micro-environments with unique characteristics.<sup>11</sup> For instance, the tumor site is weakly acidic due to the strong metabolism of tumor cells, producing excessive amounts of lactic acid and other metabolites.<sup>12</sup> Additionally, the tumor cells indicate certain reducibility since the glutathione concentration is about four times larger than in normal cells.<sup>13,14</sup> Moreover, in the tumor sites, the H<sub>2</sub>O<sub>2</sub> level is also higher than that in normal tissues because of rapid metabolism and insufficient blood supply.<sup>15</sup> Therefore, the therapeutic effect on tumours would be significantly improved, and the side effects on normal tissues would be minimized if these non-toxic and biocompatible substances in the tumor

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microenvironment could be converted into toxic substances *in situ*.

Chemodynamic therapy (CDT) generates highly toxic  $\bullet\text{OH}$  free radicals for killing tumor cells.<sup>16–18</sup> Endogenous  $\text{H}_2\text{O}_2$  molecules are directly converted into  $\bullet\text{OH}$  free radicals during the CDT process *via* Fenton or Fenton-like reactions mediated by metal ions (including  $\text{Fe}^{2+}$ ,  $\text{Cu}^+$ ,  $\text{Mn}^{2+}$ ,  $\text{Cr}^{4+}$ , and  $\text{V}^{2+}$ ).<sup>16,19–21</sup> CDT can avoid the toxic and side effects of traditional chemotherapy and resolve the limitation of light penetration due to photothermal/photodynamic therapy. Among various catalysts, ferromagnetic nanoparticles ( $\gamma\text{-Fe}_2\text{O}_3$  or  $\text{Fe}_3\text{O}_4$ ) demonstrate a dual pH-response enzyme activity *in vivo* and *in vitro*. Under neutral conditions, these iron oxide nanoparticles (IONPs) catalyze  $\text{H}_2\text{O}_2$  to decompose into non-toxic water and oxygen (catalase-like activity). However, under acidic conditions, they generate highly cytotoxic  $\bullet\text{OH}$  free radicals (peroxidase-like activity).<sup>22</sup> Therefore, IONPs can specifically generate  $\bullet\text{OH}$  free radicals within a weakly acidic tumor environment and then trigger the apoptosis of tumor cells without damaging the normal cells. However, the  $\text{H}_2\text{O}_2$  concentration is relatively low in tumor cells. Therefore, the number of  $\bullet\text{OH}$  free radicals is insufficient for killing tumor cells, thereby limiting the further application of CDT.<sup>23</sup>

The direct encapsulation of exogenous  $\text{H}_2\text{O}_2$  into tumor cells is a standard solution to the problem of low endogenous  $\text{H}_2\text{O}_2$  at the tumor site. However, it could lead to  $\text{H}_2\text{O}_2$  leakage that may damage normal tissues.<sup>6,24,25</sup> Therefore, stimulating endogenous  $\text{H}_2\text{O}_2$  in tumor cells is a preferred method. Glucose oxidase (GOx) is a common enzyme catalysing  $\beta\text{-D}$ -glucose to produce gluconic acid and  $\text{H}_2\text{O}_2$ ,<sup>26–28</sup> meeting the requirements for *in situ* generation of  $\text{H}_2\text{O}_2$ . Additionally, GOx consumes the glucose in the tumor site by blocking the energy supply of the tumor cells and leads to metabolic disorders called starvation therapy.<sup>29,30</sup> Moreover, a low concentration of endogenous  $\text{H}_2\text{O}_2$  induces the malignant transformation of normal cells, but a high concentration kills tumor cells.<sup>31</sup> Therefore, the combination of CDT with GOx-catalysed starvation therapy blocks the supply of glucose for tumor cells and increases the concentration of endogenous  $\text{H}_2\text{O}_2$  in tumor cells, which may directly kill tumor cells. However, it is more beneficial to generate highly toxic  $\bullet\text{OH}$  free radicals through the Fenton reaction to induce apoptosis in tumor cells.

Liposomes, polymer micelles, polymer nanospheres, metal-organic frameworks, and mesoporous silica nanoparticles are often observed as enzyme carriers.<sup>32–35</sup> Regenerated silk fibroin (RSF), an easily sourced and relative cheap natural polymer derived from *Bombyx mori* silkworm silk, has been extensively studied as a drug delivery carrier in cancer treatments. These silk-based carriers mainly include nanospheres<sup>36</sup> and nanofibers.<sup>37,38</sup> Previously, we developed RSF/zein nanospheres with a single central hole successfully prepared *via* a one-step method under mild conditions.<sup>39</sup> Our method was relatively simple and environmentally friendly compared to other porous materials, which utilized complicated template methods under harsh conditions.<sup>40–42</sup> The as-prepared RSF/zein nanospheres had excellent biocompatibility and biodegradability, thereby becoming the ideal enzyme carriers. This study prepared ferrous

oxide ( $\text{Fe}_3\text{O}_4$ ) and GOx dual-loaded silk fibroin/zein ( $\text{Fe}_3\text{O}_4/\text{GOx}@RSF/\text{zein}$ ) nanospheres based on this method. Moreover, the catalytic activities and the stability of the related nanospheres were thoroughly investigated. Finally, inhibiting  $\text{Fe}_3\text{O}_4/\text{GOx}@RSF/\text{zein}$  on breast cancer cells (MCF-7) through starvation therapy and CDT was evaluated.

## Experimental

### Materials

Cocoons of *B. mori* silkworm were collected from Jiangsu Province, China. Zein powder ( $\geq 97\%$ ), GOx, and glucose were procured from Sigma-Aldrich. All the other chemical reagents, such as ethanol,  $\text{Na}_2\text{CO}_3$ , LiBr, NaOH,  $\text{FeCl}_2$ , and  $\text{FeCl}_3$ , were of analytical grade and utilized without further purification.

### Preparation of RSF aqueous solution

According to an established method from our previous work, the RSF aqueous solution was prepared from *B. mori* silkworm cocoons.<sup>43</sup> In brief, the cocoons were degummed by boiling in 0.5% (w/v)  $\text{Na}_2\text{CO}_3$  for 45 min and then washing thoroughly using de-ionized water. The degummed silk was dissolved in an aqueous LiBr ( $9.3 \text{ mol L}^{-1}$ ) solution at  $45^\circ\text{C}$  for 1 h after drying at  $40^\circ\text{C}$  for 24 h. Then, the solution was dialyzed against de-ionized water with a dialysis membrane (12–14 kDa MWCO) for three days to remove the salt at room temperature. The concentration of the resulting RSF solution was measured by the weighing method.

### Preparation of RSF/zein nanospheres

Powdered zein was dissolved in 70% ethanol to obtain a solution with a  $0.625 \text{ mg mL}^{-1}$  concentration. RSF aqueous solution was diluted using de-ionized water to  $0.75 \text{ mg mL}^{-1}$ . Afterward, 5 mL of RSF aqueous solution was introduced into 2 mL of zein solution dropwise under gentle stirring for 3 min. Then, the mixture was incubated in a freezer at  $-20^\circ\text{C}$  for 24 h and thawed at room temperature. The suspension was centrifuged at 10 000 rpm for 10 min. Then, for future use, the obtained RSF/zein nanospheres were re-dispersed in de-ionized water.

### Preparation of $\text{Fe}_3\text{O}_4$ nanoparticles

4 g  $\text{FeCl}_3$  and 4.5 g  $\text{FeCl}_2$  were dissolved in 300 mL of de-ionized water. Then, the mixture was transferred into a 500 mL flask. After replacing the air with nitrogen in the flask and stabilizing it for 30 min, 15 mL of ammonia was added to the mixture. Then, it was stirred for 2 h in a nitrogen atmosphere to develop  $\text{Fe}_3\text{O}_4$  nanoparticles. The obtained  $\text{Fe}_3\text{O}_4$  nanoparticles were washed five times using de-ionized water, dried in an oven at  $40^\circ\text{C}$  for 12 h, and placed in a desiccator for further use.

### Preparation of GOx@RSF/zein nanospheres

The as-prepared RSF/zein nanospheres were dispersed in 2 mL of de-ionized water at  $1 \text{ mg mL}^{-1}$ . Then, 2 mL of GOx solution was added to the RSF/zein nanosphere dispersion, amounting to a final GOx concentration of  $1.25 \text{ mg mL}^{-1}$ . The mixture was

stirred at 100 rpm for 24 h at room temperature. The developed nanospheres were collected by centrifuging and washed thrice using de-ionized water. Finally, the as-prepared GOx@RSF/zein nanospheres were dispersed in de-ionized water and kept in a refrigerator at 4 °C for further use.

### Preparation of Fe<sub>3</sub>O<sub>4</sub>@RSF/zein nanospheres

The Fe<sub>3</sub>O<sub>4</sub> nanoparticles and zein powders were dispersed and dissolved in 70% ethanol at 0.5 mg mL<sup>-1</sup> and 1.25 mg mL<sup>-1</sup> concentrations, respectively. Then, 1 mL Fe<sub>3</sub>O<sub>4</sub> nanoparticle dispersion and 1 mL zein solution were mixed. Afterward, 5 mL RSF solution (0.75 mg mL<sup>-1</sup>) was added dropwise into the Fe<sub>3</sub>O<sub>4</sub>/zein mixture. The mixture was gently stirred for 3 min and quickly transferred to a -20 °C freezer for 24 h. The Fe<sub>3</sub>O<sub>4</sub>@RSF/zein nanospheres were collected with a magnet after thawing at room temperature, washing with 20% ethanol thrice and de-ionized water thrice, and dispersed in de-ionized water.

### Preparation of Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres

2 mL GOx solution (2.5 mg mL<sup>-1</sup>) was mixed with a 2 mL Fe<sub>3</sub>O<sub>4</sub>@RSF/zein dispersion (1 mg mL<sup>-1</sup>) at room temperature by gently stirring for 24 h. The precipitate was collected using a magnet and washed with de-ionized water thrice. Finally, the magnetic Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres were dispersed in de-ionized water and placed in a refrigerator at 4 °C.

### Morphology observations

All the nanosphere solutions were diluted to 100 µg mL<sup>-1</sup> before observation under a scanning electroprobe microscope (SEM). SEM images were obtained with a Hitachi S-4800 high-resolution SEM at 1 kV. The nanosphere solutions were diluted to 50 µg mL<sup>-1</sup> and dropped on carbon-coated copper grids before observing with transmission electroprobe microscopy (TEM). TEM images were obtained with a Tecnai G2 TEM at 200 kV.

### Size and zeta potential

A Zetasizer Nano ZS 90 (Malvern Inst. Ltd, UK) was used to analyze the size and zeta potential of all nanospheres. The as-prepared nanosphere solution was diluted before each analysis to satisfy the requirements of the equipment and then filtered over an 800 nm filter to remove the dust.

### X-ray diffraction

Fe<sub>3</sub>O<sub>4</sub> nanoparticles, RSF/zein nanospheres, and Fe<sub>3</sub>O<sub>4</sub>@RSF/zein nanospheres were tested with an X'Pert Pro X-ray powder diffractometer (PANalytical, Netherlands) using CuKα radiation. The working voltage was 40 kV, the current was 40 mA, and the scanning range was between 10–80°.

### Thermogravimetric analysis

A Pyris 1 thermogravimetric analyzer (PerkinElmer, USA) was utilized to test RSF/zein and Fe<sub>3</sub>O<sub>4</sub>@RSF/zein nanospheres under an air atmosphere. The heating rate was 10 °C min<sup>-1</sup>, and the temperature range was 50–800 °C.

### The catalytic activity of GOx

The pristine GOx was dissolved, and the GOx@RSF/zein nanospheres were dispersed in de-ionized water with a 50 µg mL<sup>-1</sup> GOx concentration. Then, 1 mL GOx solution or the GOx@RSF/zein nanosphere dispersion was mixed with 4 mL glucose solution under different concentrations (100, 200, 400, 800, 1000, and 1250 µg mL<sup>-1</sup>) at room temperature. The pH of the GOx solution or the GOx@RSF/zein nanosphere dispersion and the H<sub>2</sub>O<sub>2</sub> concentration was monitored after 1 h. H<sub>2</sub>O<sub>2</sub> production in the solution was characterized using classical colorimetry.<sup>41</sup>

### Stability of GOx@RSF/zein nanospheres

To study the long-time stability of GOx@RSF/zein nanospheres, the dispersion was stored in a refrigerator at 4 °C for 1, 7, 15, 30, and 60 days before testing. To determine the catalytic activity of GOx@RSF/zein nanospheres stored for different periods of time, the abovementioned method was applied at room temperature. After 1 h of reaction with glucose, the GOx@RSF/zein nanospheres were obtained and re-dispersed in a fresh glucose solution to react for another 1 h to evaluate the cycle stability. The process was repeated five times.

### Degradation ability test

GOx@RSF/zein, Fe<sub>3</sub>O<sub>4</sub>@RSF/zein, and Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanosphere dispersions were added into the glucose solution. Then, methylene blue solution was added to prepare the final glucose concentration of 1 mg mL<sup>-1</sup> and the methylene blue concentration of 10 µg mL<sup>-1</sup>. The absorbance was used to determine the degradation of methylene blue at 644 nm, which was measured using a UV-vis spectrophotometer (Hitachi U-2910, Japan).

### In vitro cytotoxicity

The *in vitro* cytotoxicity was determined as the cell viability using the CCK-8 assay. Human breast cancer MCF-7 cells were grown in DMEM culture medium with 10% fetal bovine serum (FBS) and 1% penicillin-streptomycin. Next, the cells were seeded into 96-well plates at a density of 1 × 10<sup>4</sup> cells per well and were grown for 24 h at 37 °C. Then, varying concentrations of GOx@RSF/zein, Fe<sub>3</sub>O<sub>4</sub>@RSF/zein, and Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres were introduced into the medium solutions. The cells were cultured for another 24 h and washed with PBS. Subsequently, a medium with 10% CCK-8 solution was added to each well and further incubated for 2 h. The absorbance of each well was determined at 450 nm using a microplate reader (Bio-Tek, ELx800, USA). The relative cell viability (%) was calculated as follows:

$$\text{Cell viability (\%)} = \frac{[A]_{\text{test}}}{[A]_{\text{control}}} \times 100\%$$

where [A]<sub>test</sub> is the absorbance of the test sample and [A]<sub>control</sub> is the absorbance of the control sample incubated without any drugs or nanospheres within the medium solution.

### In vitro cellular uptake of nanospheres

The cellular uptake of RSF/zein, GOx@RSF/zein, Fe<sub>3</sub>O<sub>4</sub>@RSF/zein, and Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres into MCF-7 cells

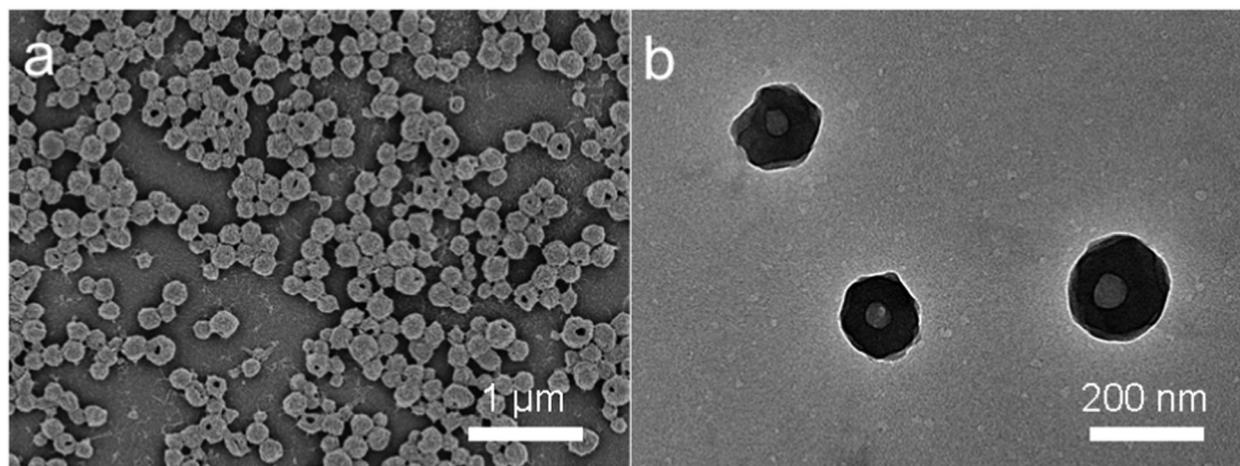


Fig. 1 Morphology of the GOx@RSF/zein nanospheres. (a) SEM and (b) TEM.

was observed under a C2 confocal laser scanning microscope (CLSM) (Nikon, Japan). The MCF-7 cells were incubated in a confocal dish at  $2 \times 10^5$  cells per dish density for 24 h at 37 °C. The cells were washed thrice using PBS and then fixed with 4% (w/w) glutaraldehyde for 15 min inside an ice bath after incubating with RITC-labelled nanospheres for an additional 4 h. After washing another three times with PBS, DAPI was

added to stain the nucleus of the cells for 10 min. The excess DAPI was removed by washing with PBS. Then, 2 mL PBS was added to regulate the cell morphology.

#### Production of reactive oxygen species (ROS) in cells

MCF-7 cells were incubated at  $2 \times 10^5$  cells per dish density in a confocal dish for 24 h at 37 °C. RSF/zein, GOx@RSF/zein,

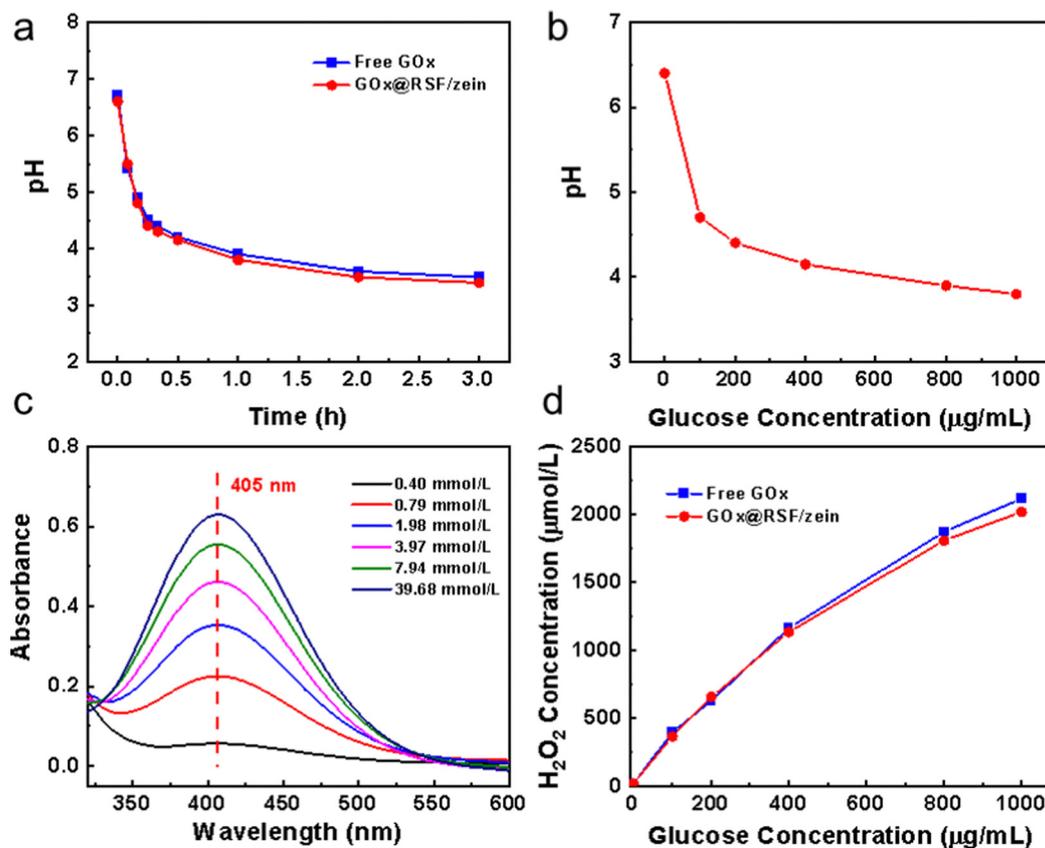


Fig. 2 The performance of GOx@RSF/zein nanospheres in the catalytic reaction of glucose solution. (a) pH change with reaction time (glucose solution: 1 mg mL<sup>-1</sup>), (b) pH changes with glucose concentration (reaction time: 1 h), (c) UV-vis spectra of Ti(IV)O<sub>2</sub>SO<sub>4</sub> solution under different H<sub>2</sub>O<sub>2</sub> concentrations, and (d) H<sub>2</sub>O<sub>2</sub> generated in glucose solution under different concentrations (reaction time: 1 h).

$\text{Fe}_3\text{O}_4@\text{RSF}/\text{zein}$ , and  $\text{Fe}_3\text{O}_4/\text{GOx}@\text{RSF}/\text{zein}$  nanospheres were diluted using serum-free media. Then, 2 mL of nanosphere solution were co-incubated with MCF-7 cells for 4 h. After rinsing thrice with PBS, 2',7'-dichlorofluorescein diacetate (DCFH-DA) fluorescent probes were added to incubate for another 1 h. The excess DCFH-DA was washed using PBS, and then 2 mL PBS was added to regulate the cell morphology. The green fluorescence inside the cells was observed with a C2 confocal laser microscope (Nikon, Japan)

## Results and discussion

### Preparation and catalytic performance of $\text{GOx}@\text{RSF}/\text{zein}$ nanospheres

In our previous work, RSF/zein nanospheres with excellent biocompatibility were successfully prepared *via* a one-step method.<sup>39</sup> They had a narrow size distribution and good dispersibility within water. More importantly, the RSF/zein nanosphere had a unique structure with a single central hole of 26 nm average size. RSF/zein nanospheres were more suitable for loading macromolecules, including nucleic acids and proteins compared with the traditional mesoporous silica nanoparticles (with a pore size of about 5 nm).<sup>42</sup> Therefore, GOx, with a size of  $6.0 \times 5.2 \times 7.7$  nm, could be a good

candidate for loading, with a capacity as high as 7.4%. Fig. 1 shows the SEM and TEM images of  $\text{GOx}@\text{RSF}/\text{zein}$  nanospheres, which seldom indicate morphological changes compared with the pristine RSF/zein nanospheres.<sup>39</sup> Additionally, the size of the nanospheres did not change significantly before and after GOx loading. However, the zeta potential was increased from  $-26.2$  to  $-22.1$  mV (Table S1, ESI<sup>†</sup>). This established the successful loading of GOx, consistent with a literature report.<sup>6</sup>

The glucose decomposes into  $\text{H}_2\text{O}_2$  and gluconic acid due to GOx catalysis.<sup>45</sup> Therefore, the glucose solution pH decreases during the reaction process. Fig. 2a shows that the pH dropped rapidly from 6.7 to 4.8 within 10 min after adding the  $\text{GOx}@\text{RSF}/\text{zein}$  nanospheres to the glucose solution. The reaction reached equilibrium at about 3 h, and the final pH was approximately 3.4. Meanwhile, the catalytic behaviour of the  $\text{GOx}@\text{RSF}/\text{zein}$  nanospheres was similar to the pristine GOx, indicating that its catalytic activity is unaffected by the loading of GOx in the RSF/zein nanospheres. In addition, the pH of the glucose solution decreased more significantly with the increase in glucose concentration. Fig. 2b indicates that, after 1 h of reaction, the pH of the solution decreased from 4.7 to 3.8 when the glucose concentration was elevated from 100 to  $1000 \mu\text{g mL}^{-1}$ .

The  $\text{H}_2\text{O}_2$  concentration in the glucose solution can be monitored by the colour change (Fig. 2c) since titanium oxysulfate

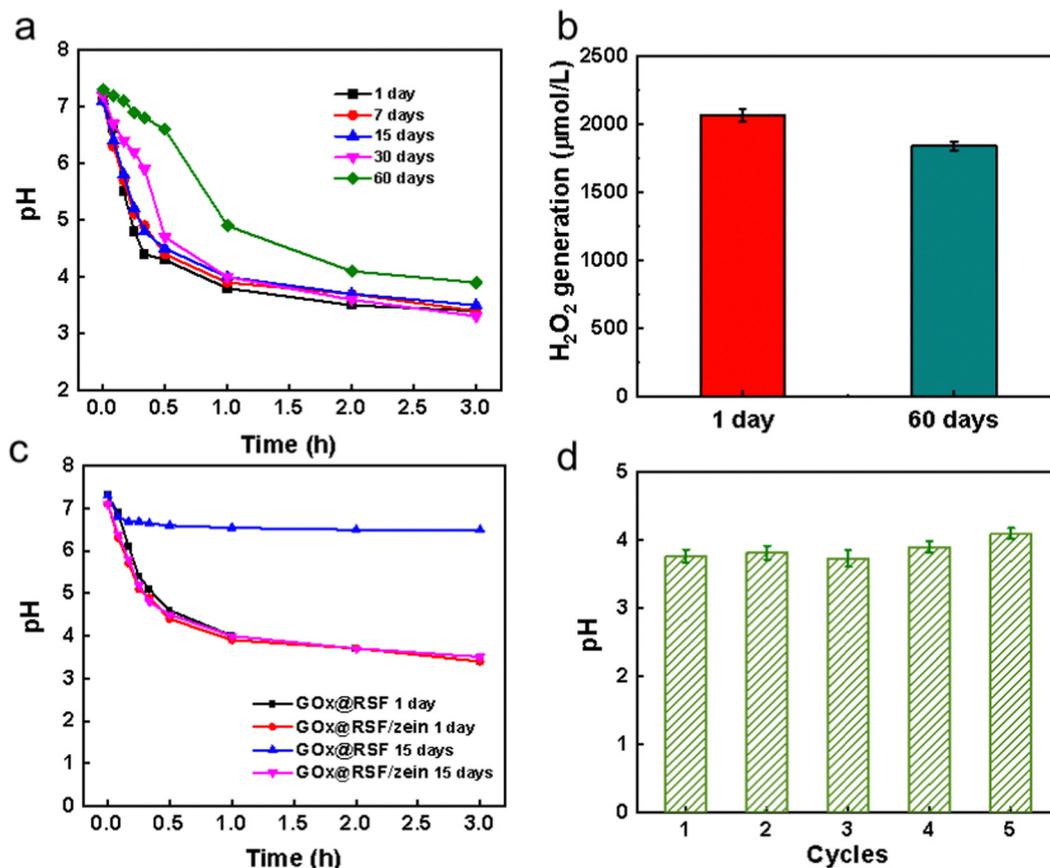


Fig. 3 The stability of  $\text{GOx}@\text{RSF}/\text{zein}$  nanospheres. (a) pH change with the reaction time after a different storage period (glucose solution:  $1 \text{ mg mL}^{-1}$ ), (b)  $\text{H}_2\text{O}_2$  generation after 1 day and 60 days of storage, (c) pH change compared with  $\text{GOx}@\text{RSF}$  nanospheres, and (d) pH value during different cycles.

forms a yellow precipitation titanium peroxide complex ( $\text{Ti(IV)O}_2^{2+}$ ) with the presence of  $\text{H}_2\text{O}_2$ .<sup>44</sup> Fig. 2d represents the  $\text{H}_2\text{O}_2$  concentration generated in the glucose solution after an hour reaction using GOx. It shows that the  $\text{H}_2\text{O}_2$  concentration enhanced almost linearly with the increase in glucose concentration within the catalytic system. Moreover, there is minimal difference between the  $\text{H}_2\text{O}_2$  concentration–glucose concentration curve from pristine GOx and GOx@RSF/zein nanospheres. Therefore, it is further proved that the GOx loaded in the RSF/zein nanospheres has similar catalytic activity as its original form.

The stability of GOx@RSF/zein nanospheres was evaluated by monitoring their catalytic activities after storing them for 1, 7, 15, 30, and 60 days at 4 °C. Fig. 3a revealed the pH changes at different time points when GOx@RSF/zein nanospheres were introduced in 1 mg mL<sup>-1</sup> glucose solution. The pH change of these nanospheres stored for 30 days achieved equilibrium at about pH = 3.5 after following the same trend. Only the nanospheres kept for 60 days had a smaller pH change at each time point than other storage times, and the equilibrium pH was elevated to 3.8.  $\text{H}_2\text{O}_2$  generated by the GOx@RSF/zein nanospheres showed the same trend as the pH change. In the glucose solution, the  $\text{H}_2\text{O}_2$  concentration was about 2000  $\mu\text{mol L}^{-1}$  for those GOx@RSF/zein nanospheres stored for 1 to 30 days. However, it decreased to about 1800  $\mu\text{mol L}^{-1}$  after 60 days of

storage while maintaining 90% of catalytic activity (Fig. 3b). This clearly indicates that the RSF/zein nanospheres kept the activity of GOx for at least one month. After a longer time, for example two months, the GOx activity could be lost a little due to the partial degradation of the RSF/zein nanospheres.

The stability of the GOx catalytic activity was also compared after loading in pristine RSF nanospheres and RSF/zein nanospheres. As depicted in Fig. 3c, GOx@RSF and GOx@RSF/zein nanospheres demonstrated similar catalytic activity on the first day. Under these circumstances, the pH of the glucose solution reduced to 3.4 after 3 h of reaction. However, the GOx@RSF/zein nanospheres still showed similar excellent catalytic activity after storing for 15 days, but the GOx@RSF nanospheres almost lost their catalytic activity. This suggests that the unique structure of the RSF/zein nanospheres can protect the GOx from inactivation.

The cyclic stability test of the GOx@RSF/zein nanospheres was performed by collecting the nanospheres in a glucose solution after reacting for 1 h. Later, the nanospheres were transferred into a fresh glucose solution for another hour of catalysis. Fig. 3d indicates no significant reduction in catalytic activity after five cycles. This further suggested that the RSF/zein nanospheres protected the catalytic activity of GOx.

Rapid tumor site metabolism leads to a special tumor microenvironment. The fast proliferation of tumor cells

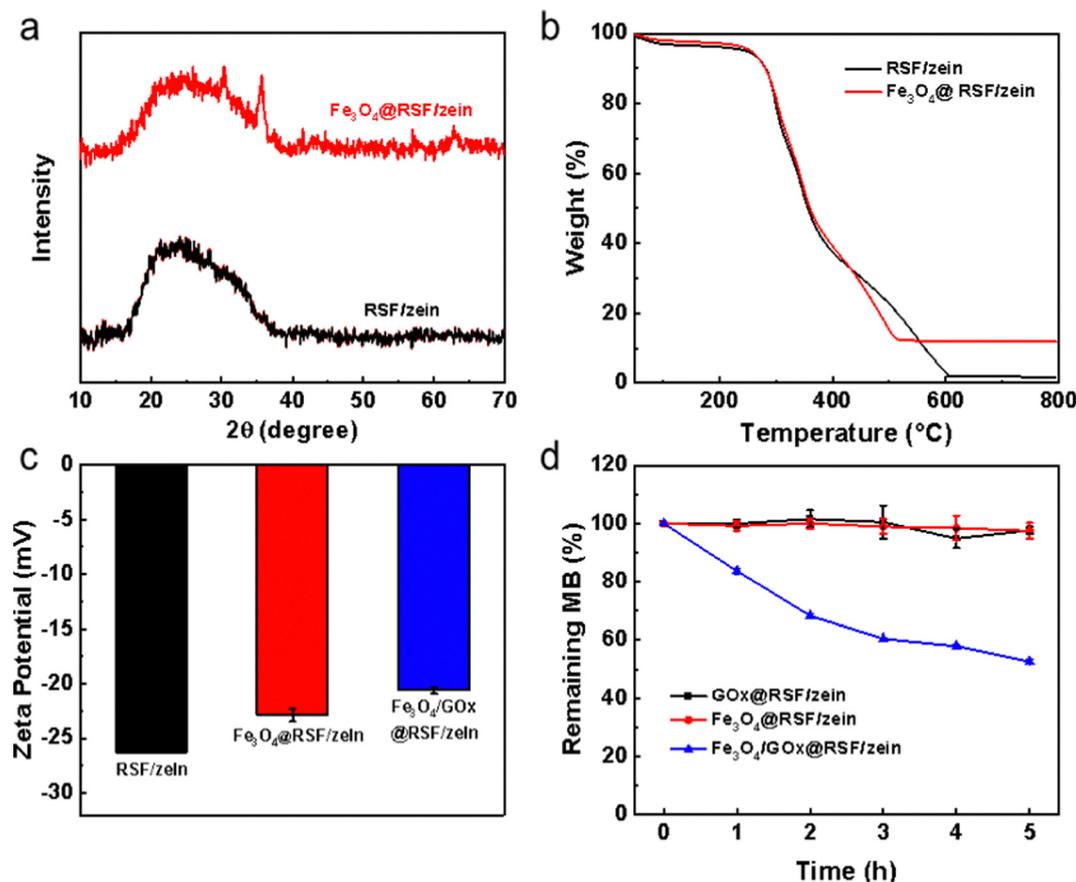


Fig. 4 The characterization of  $\text{Fe}_3\text{O}_4$ @RSF/zein and  $\text{Fe}_3\text{O}_4$ /GOx@RSF/zein nanospheres. (a) XRD pattern, (b) TGA curve, (c) zeta potential, and (d) degradation of methylene blue.

depends on the extensively constructed blood vessels in the tumor to provide nutrition for them, with glucose being an essential nutrient.<sup>46</sup> Based on the Warburg effect, tumor cell proliferation primarily depends on aerobic glycolysis, making tumor cells very sensitive to glucose alteration.<sup>47</sup> Therefore, changing the glucose metabolism pathway inside the tumor cells is a safe and efficient strategy.<sup>48</sup> The above results demonstrated that RSF/zein nanospheres effectively protect the catalytic activity of GOx. The active GOx consumes glucose at tumor sites, cuts off the nutrient supply, and generates a high  $\text{H}_2\text{O}_2$  concentration. Therefore, it promotes the subsequent Fenton reaction to cause tumor cell apoptosis.

### Construction of $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$ nanospheres

The X-ray diffraction pattern of the  $\text{Fe}_3\text{O}_4$  nanoparticles is demonstrated in Fig. S1a (ESI<sup>†</sup>). The prominent diffraction peaks are at  $30.1^\circ$ ,  $35.5^\circ$ ,  $43.1^\circ$ ,  $57.0^\circ$ , and  $62.6^\circ$ , corresponding to the (220), (311), (400), (511), and (440) planes, respectively. Thus, the product is a typical inverse cubic spinel  $\text{Fe}_3\text{O}_4$ ,<sup>49</sup> with an average size of 28 nm and a narrow distribution based on the DLS measurement (Fig. S1b, ESI<sup>†</sup>).

During the preparation process of the RSF nanospheres,  $\text{Fe}_3\text{O}_4$  nanoparticles can be loaded inside these nanospheres by

directly adding them.<sup>50</sup>  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanospheres were successfully prepared following this method. Fig. 4a shows the X-ray diffraction pattern before and after loading the  $\text{Fe}_3\text{O}_4$  in RSF/zein nanospheres. The characteristic diffraction peaks of  $\text{Fe}_3\text{O}_4$ , including  $35.5^\circ$  and  $62.6^\circ$ , were seen in  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanospheres. In addition, the mass fraction of  $\text{Fe}_3\text{O}_4$  loaded in the RSF/zein nanospheres was about 10% through TGA analysis (Fig. 4b). The size of  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  and  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres could be a little larger than the pristine RSF/zein nanospheres. However, the average sizes were around 300 nm (Fig. S1c, ESI<sup>†</sup>). Moreover, the zeta potential of the  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanospheres increased from  $-26.2$  to  $-22.9$  mV compared with the pristine RSF/zein nanospheres, which further enhanced to  $-20.6$  mV for  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres (Fig. 4c). This is consistent with the literature report,<sup>6</sup> confirming the successful loading of  $\text{Fe}_3\text{O}_4$  and GOx within the RSF/zein nanospheres at the same time.

$\text{Fe}_3\text{O}_4$  nanoparticles can react with  $\text{H}_2\text{O}_2$  to generate  $\cdot\text{OH}$  free radicals through the Fenton reaction. Therefore, methylene blue was chosen as an indicator to explore the *in vitro* generation of  $\cdot\text{OH}$  free radicals.<sup>19</sup> Only the  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres significantly degraded the methylene blue

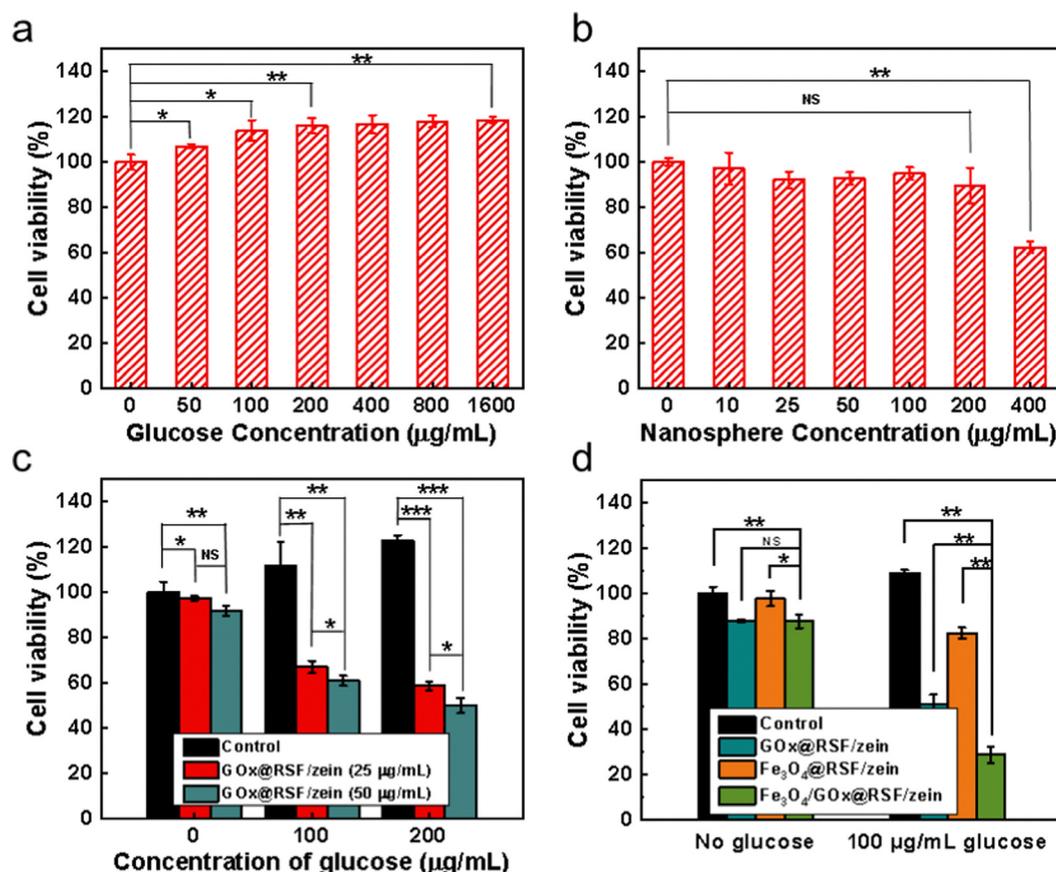


Fig. 5 The *in vitro* cytotoxicity of different RSF/zein nanospheres. (a) The effect of glucose concentration on the cell viability without nanospheres, (b) the effect of  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanosphere concentration on cell viability, (c) the effect of glucose concentration on cell viability under different concentrations of  $\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres, and (d) the comparison of cell viability among  $\text{GOx}@/\text{RSF}/\text{zein}$ ,  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$ , and  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres ( $50 \mu\text{g mL}^{-1}$ ). \* $p < 0.05$ , \*\* $p < 0.01$ , \*\*\* $p < 0.001$ .

(Fig. 4d). This indicated that only in  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres, GOx catalysed glucose to generate  $\text{H}_2\text{O}_2$ , then the  $\text{H}_2\text{O}_2$  reacted with  $\text{Fe}_3\text{O}_4$  to generate  $\bullet\text{OH}$  free radicals, and thereby degraded methylene blue.

#### *In vitro* cytotoxicity of $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$ nanospheres

Glucose is the primary energy source for tumor growth, and its concentration increase enhanced MCF-7 cell proliferation (Fig. 5a). Meanwhile,  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanospheres revealed no toxicity toward the cells at the concentration range between  $10\text{--}200\ \mu\text{g mL}^{-1}$  (Fig. 5b).  $\text{Fe}_3\text{O}_4$  toxicity in the cells is mainly derived from the Fenton reaction with endogenous  $\text{H}_2\text{O}_2$  generating cytotoxic  $\bullet\text{OH}$  free radicals. However, the effect of  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanospheres on cytotoxicity is limited due to the low concentration of endogenous  $\text{H}_2\text{O}_2$  in the cells. Thus, it almost does not depend on the  $\text{Fe}_3\text{O}_4$  concentration. Therefore, the elevated cytotoxicity observed at a high  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanosphere concentration ( $400\ \mu\text{g mL}^{-1}$ ) was primarily attributed to the high concentration of the nanospheres rather than the generation of more  $\bullet\text{OH}$  free radicals.

The starvation therapy depending on GOx, primarily relies on consuming glucose and generating a high concentration of  $\text{H}_2\text{O}_2$  to induce cell apoptosis. It has a stronger anticancer performance than traditional starvation therapy while cutting off the glucose supply.<sup>29</sup> Fig. 5c indicates the cytotoxicity of  $\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres with two different concentrations. It shows the cell viability increases with the increase in glucose concentration if there are no  $\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres. However, the cell viability significantly decreases with the increase in glucose concentration if there are  $\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres. This decrease was observed with the increase in  $\text{GOx}@/\text{RSF}/\text{zein}$  nanosphere concentration, confirming the curative effect of the  $\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres.

Finally, the effect of  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres on cell viability was determined and compared with  $\text{GOx}@/\text{RSF}/\text{zein}$  and  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanospheres (Fig. 5d). This revealed the synergistic effect of starvation therapy and CDT for  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres to kill or inhibit cancer cells.

The silk fibroin was labelled with RITC, so the nanospheres show red fluorescence. Then, they were used to determine whether the  $\text{GOx}@/\text{RSF}/\text{zein}$ ,  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$ , and  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres could enter MCF-7 cells or not. The cell nucleus was further stained using DAPI after co-incubating the RITC-labelled nanospheres with MCF-7 cells for 4 h. Fig. 6 shows that  $\text{RSF}/\text{zein}$ ,  $\text{GOx}@/\text{RSF}/\text{zein}$ ,  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$ , and  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres enter the cell very well. In addition, a series of nine-step z-stack confocal fluorescence images of MCF-7 cell after incubating with  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres was shown in Fig. S2 (ESI<sup>†</sup>), which further indicated that  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres entered MCF-7 cells.

Finally, the catalytic performance of  $\text{GOx}@/\text{RSF}/\text{zein}$ ,  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$ , and  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres in MCF-7 cells was evaluated. The nanospheres were stained with a 2',7'-dichlorofluorescein yellow diacetate (DCFH-DA) fluorescent probe (DCFH-DA is oxidized to 2',7'-dichlorofluorescein

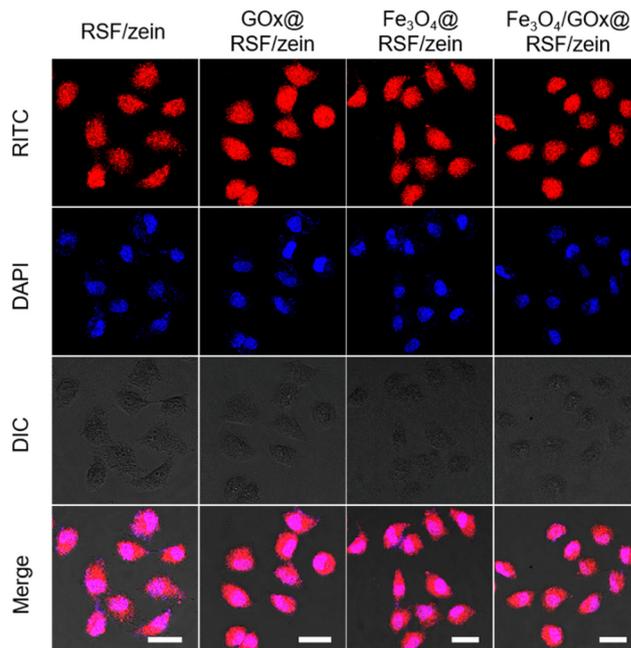


Fig. 6 The cellular uptake of  $\text{RSF}/\text{zein}$ ,  $\text{GOx}@/\text{RSF}/\text{zein}$ ,  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$ , and  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres within MCF-7 cells after incubating for 4 h at  $37\ ^\circ\text{C}$ . Scale bar:  $20\ \mu\text{m}$ .

(DCF) with the ROS, depicting green fluorescence<sup>51</sup>) after co-incubating them with MCF-7 cells for 4 h. As shown in Fig. 7, only weak green fluorescence was observed in the cells with  $\text{RSF}/\text{zein}$  and  $\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres. This mainly results from a small number of ROS due to cell metabolism. The green fluorescence was slightly enhanced in the cells with  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$  nanospheres, indicating the generation of ROS was slightly increased. The green fluorescence was significantly enhanced in those cells with  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres. Therefore, a

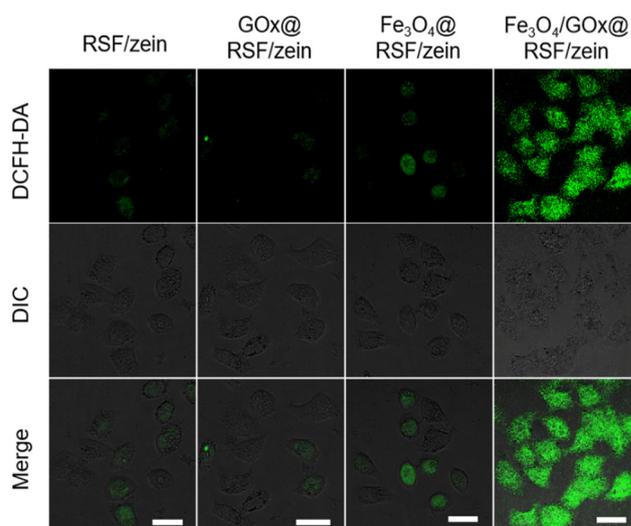


Fig. 7 The intracellular  $\bullet\text{OH}$  detection in MCF-7 cells after incubating using  $\text{RSF}/\text{zein}$ ,  $\text{GOx}@/\text{RSF}/\text{zein}$ ,  $\text{Fe}_3\text{O}_4@/\text{RSF}/\text{zein}$ , and  $\text{Fe}_3\text{O}_4/\text{GOx}@/\text{RSF}/\text{zein}$  nanospheres. Scale bar:  $20\ \mu\text{m}$ .

large number of ROS were generated in the cells because of the synergistic effect of GOx and Fe<sub>3</sub>O<sub>4</sub>. The ROS observed in the cells were assumed mainly to be the •OH free radicals according to the experiment results shown above, but this will be further confirmed in future studies. These results indicate that GOx and Fe<sub>3</sub>O<sub>4</sub> undergo the same outstanding catalytic function in a complicated cellular environment, first by decomposing glucose to generate H<sub>2</sub>O<sub>2</sub> and then developing the highly cytotoxic •OH free radicals through the Fenton reaction.

## Conclusions

The current work introduced Fe<sub>3</sub>O<sub>4</sub> nanoparticles and GOx to the RSF/zein nanospheres that were previously developed in our laboratory for preparing dual-loaded animal protein/plant protein composite nanospheres. Immobilizing GOx in RSF/zein nanospheres did not alter its catalytic activity. The catalytic activity was maintained after the nanospheres were stored at 4 °C for at least one month. Additionally, the catalytic activity of GOx loaded in the RSF/zein nanospheres remained nearly the same after five catalytic reaction cycles. This suggests that the RSF/zein nanosphere could be an efficient and safe nanocarrier for GOx. Methylene blue was significantly degraded within 5 h after placing the dual-loaded nanospheres (*i.e.*, Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres) in glucose/methylene blue solution. In contrast, the single-loaded nanospheres (either GOx@RSF/zein or Fe<sub>3</sub>O<sub>4</sub>@RSF/zein nanospheres) did not show the same phenomenon, indicating the co-existence of Fe<sub>3</sub>O<sub>4</sub> and GOx in the RSF/zein nanosphere led to a continuous catalytic reaction. First, GOx decomposed the glucose into H<sub>2</sub>O<sub>2</sub> and gluconic acid and reduced the pH of the glucose solution. Then, Fe<sub>3</sub>O<sub>4</sub> reacted with H<sub>2</sub>O<sub>2</sub> to generate highly cytotoxic •OH free radicals under acidic conditions through the Fenton reaction.

After co-loading GOx and Fe<sub>3</sub>O<sub>4</sub>, they indicated significant cytotoxicity toward cancer cells (MCF-7 cells) despite the excellent biocompatibility of pristine RSF/zein nanospheres. Additionally, the antiproliferative activity of Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres towards MCF-7 cells was higher than that of either GOx@RSF/zein or Fe<sub>3</sub>O<sub>4</sub>@RSF/zein nanospheres. Therefore, this indicates that the combination of starvation therapy and CDT has a synergistic effect during cancer treatment. Finally, the Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres were taken by MCF-7 cells, and the highly cytotoxic •OH free radicals existed within the cells. Therefore, the combination of starvation therapy and CDT provided by the Fe<sub>3</sub>O<sub>4</sub>/GOx@RSF/zein nanospheres in the tumor microenvironment exhibits a significant synergistic anticancer effect with less toxic side effects to normal tissues. Therefore, it shows excellent promise for utilization in further clinical applications.

## Conflicts of interest

There are no conflicts to declare.

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