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Characterizing catalyst function and transformations in the plasma reduction of CO₂ on atomic layer deposition-synthesized catalysts†

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The enhancement of CO₂ reduction in atmospheric-pressure, non-thermal plasma has been shown using a variety of catalyst systems with ranging composition, particle sizes, and morphologies. Improvements in CO₂ conversion can be attained by choice of catalyst material. However, inhomogeneity in the material distribution arising from the synthesis affects the catalytically active surface area and dielectric environment that modulates the plasma properties near the catalyst. Atomic layer deposition (ALD) can be used to control the composition of ultra-thin layers on support materials. We used ALD to synthesize metal oxide catalyst coatings on high surface area supports. We found that TiO₂ achieved significantly higher yields of CO₂ conversion (to CO and O₂) at low reactor power compared to ZnO or Al₂O₃, materials commonly used as a support for other catalysts. We also observed an unexpected increase in the catalytic activity on ZnO with increasing power. The results here suggest that ALD can unambiguously isolate the catalytic effects of materials in plasma reactors.

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Introduction

The steady rise of atmospheric CO₂ due to fossil fuel use has resulted in increasing global average temperatures.^{1–3} Due to environmental concerns and developing economic incentives, a number of technologies have been developed for transforming captured CO₂ into more useful molecules as a feedstock for other chemical processes.^{4–7} One promising example is the use of non-thermal plasmas to convert CO₂ into CO and O₂, with the product CO of use to other industrial or synthetic applications.^{8–14} The dielectric barrier discharge (DBD) reactor is commonly used for this reaction, as it is relatively straightforward to construct and is capable of producing a non-thermal plasma with relatively high-energy electrons (1–10 eV) and low bulk gas temperatures.^{9,14–16} These high energy electrons are capable of directly splitting CO₂ into CO and O₂. Particulate heterogeneous catalysts composed of metal oxide materials such as ZnO, Al₂O₃, or multi-metal oxides are often packed inside the discharge region of the reactor to improve both

conversion rates and efficiency.^{9,17,18} The catalyst material is commonly formed by distributing chemically-synthesized catalyst particles either as a packed bed or deposited onto an inert support material inside the plasma volume. Particles are generally deposited onto the support by capillary solvent adsorption or vapor deposition techniques.^{19–22} The structure of the support and catalyst has been shown to influence the overall CO₂ conversion by affecting the true active catalytic surface area, the flow dynamics of reactant gases, and the electric field inside the discharge region.^{9,23–25} Specifically, the effects of macroscopic catalyst loading and morphology on plasma-catalyst interactions have been explored in other works. For example, increasing support packing density has been shown to affect the primary discharge mode of the reactor, decreasing the selectivity towards CO during plasma-catalytic dry reforming of methane (DRM).^{23–25} The morphology of the catalyst itself has also been shown to be important, with nanoparticle Ni dendrites and flakes on alumina support having been shown to be more active for plasma-catalytic DRM than identical weight loadings of Ni nanospheres.²⁶ Clearly, to make direct comparisons of different catalysts for plasma CO₂ reduction requires consideration of the synthesis and decoration method, even for identical chemical composition.

While the use of particulate catalysts may allow for convenient control during synthesis, the heterogeneity of the distribution can have significant effects on the process. It is difficult to separate catalytic effects from those derived from

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mass transport or interactions with the plasma in the reactor. Atomic layer deposition (ALD) is a self-limiting, highly-controlled growth process that is capable of depositing conformal films with atomic control over layer thickness and composition.^{27–31} ALD has also been shown to be capable of conformally coating the nanopores of support materials without changing the overall porosity or microstructure of the support.^{32,33} This control, along with the low-pressure deposition environment, allows for the coating of high aspect ratio or geometrically complex features, such as microwires or foams, without the line of sight restrictions imposed by other coating techniques such as sputtering or flame-spray.^{34–37} ALD has been used to synthesize electrocatalysts or thermal catalysts on supports for CO₂ reduction.^{38,39} With respect to plasma catalysis, ALD has primarily been used as a method for applying catalysts developed for pollutant degradation and non-oxidative methane coupling.^{38,40}

In this work, we characterize the activity of ALD-synthesized thin films of metal oxide catalysts on an aluminosilicate support for catalyzing CO₂ reduction in a DBD plasma reactor. This approach regularizes the structure of the support and isolates the catalytic effects of the nanoscale, conformal coating. We used ALD to prepare three metal oxide catalysts of interest for CO₂ reduction in a DBD plasma reactor to produce CO and O₂. TiO₂ has been investigated as a plasma CO₂ reduction catalyst, and is therefore a well-known catalyst around which to develop this technique.^{41–46} Al₂O₃ was chosen because it has been shown to be much less active for plasma CO₂ reduction than other metal oxides and has often been used as a support material for other catalysts.^{9,17,47,48} ZnO is typically used in DBD plasma reactions as a support material for other co-catalysts such as Pd or Cu for CO₂ hydrogenation reactions.^{38,49–51} While both TiO₂ and Al₂O₃ behaved as expected, we observed a significant increase in the activity of ZnO with increasing reactor power. Regularizing the catalyst structure shows that this transition is only explained by a transition in the behavior between the ZnO surface and the plasma. There is significant catalytic behavior from a material that is conventionally considered a support material. More generally, this work shows that ALD can be used to isolate catalytic activity from other effects (mass transport, plasma variability near dielectric interfaces) in complex plasma reduction reactions.

Materials and methods

ALD catalyst preparation

The support used as a substrate for the deposition of catalysts was a commercial aluminum-silicate wool (Lynn Manufacturing, USA) designed to withstand high temperatures. The support was comprised of fibers that were roughly 5–8 μm in diameter as determined by optical microscopy and was used as received from the manufacturer. Sections of the support (1.010 g, 60 × 60 × 7 mm) were cut and cleaned using a successive solvent wash of acetone (99%,

EMD Millipore Corp), methanol (HPLC grade, VWR Analytical), isopropanol (HPLC grade, VWR Analytical) and water (HPLC grade, VWR Analytical) respectively, all of which were used as received. Once cleaned, the supports were dried at 180 °C for 1 hour before a final treatment in a UV-ozone cleaner for 30 minutes. The supports were then used directly for preparing the ALD-synthesized catalysts.

The ALD layers were grown in a commercial ALD reactor (GEMstar XT; Arradance Inc.) by a sequence of exposures of an organometallic precursor and water as an oxygen source. The precursors were contained in a vial sealed with a metal gasket: trimethylaluminum (TMA, 98%; Strem, Inc.) for Al₂O₃, diethylzinc (DEZ, 95%; Strem, Inc.) for ZnO, and tetrakis(dimethylamido)titanium (TDMAT, 99%; Strem, Inc.) for TiO₂. The specific details of each ALD cycle for the growth of these oxides are provided in the ESI.† Each heterogeneous catalyst was composed of an ALD-synthesized layer conformally coating the support. The ALD coatings began with a 50-cycle base layer of Al₂O₃ onto the support to provide a consistent interface between each catalyst and the support. The catalyst layer was deposited over this base layer: 200 cycles of TDMAT/H₂O to prepare TiO₂, 58 cycles of DEZ/H₂O to prepare ZnO, or 93 additional cycles of TMA/H₂O to prepare Al₂O₃. These layer structures were chosen to grow catalyst layers of sufficient thickness to behave like the bulk compound yet add negligible thickness to the fibers. After the deposition was completed, the samples were removed from the ALD reactor and annealed in air at 500 °C for three hours (after a 3 °C min⁻¹ ramp from room temperature). This was done to remove any remaining unreacted precursor detritus left behind from the ALD process and to controllably carry out calcination of the films prior to plasma exposure, both of which would affect the activity of the catalysts.

Reactor configuration and operation

We developed the reactor setup shown in Fig. 1a and b to compare the performance of ALD-prepared plasma catalysts. The catalysts (ALD-coated fiber supports) were loaded into the plasma volume. This volume was established by the length of the exterior aluminum foil electrode (20 μm thick). The length of the exterior electrode was 55 mm and wrapped three times around the tube. It was connected to the transformer by a metal hose clamp which also acted to mechanically fix the electrode to the quartz tube (25 mm outer diameter). The catalyst was fixed in place by packing it between the inner stainless-steel electrode (15.8 mm outer diameter) and the quartz wall of the tube (1.6 mm thick). The radial discharge gap was fixed at roughly 3.0 mm between the interior electrode and dielectric barrier, both of which were cleaned and polished between samples. An AC signal (12–20 kV, 19 kHz) was provided between the interior and exterior electrodes by an AC transformer (Allanson SS1235OX). The input power provided to the transformer (denoted input power) was monitored by a power meter measuring the power into the transformer. The plasma input





Fig. 1 (a) A schematic of the DBD plasma reactor used to compare the catalytic activity of ALD-coated supports. The end caps of the reactor are omitted for clarity. (b) A photograph of the discharge region of the reactor during operation.

power could therefore be set using a variable autotransformer for each experiment to study the influence of plasma input power on CO₂ conversion and catalyst stability. The reactor was not cooled during any experiments and thermal management was provided by controlling the plasma ‘on’ time, during which the exterior of the reactor never exceeded 250 °C (Fig. S1†). Input CO₂ flow was provided by a mass flow controller (SmartTrak 50; Sierra). The reactor was continuously fed CO₂ (Airgas 99.999%) at a rate of 75 sccm and maintained at 1 atm. The reactor gas passed through the discharge region, then into an on-line gas chromatograph (6180-GC; SRI) containing two 1.1 mL auto-sampling loops. One loop delivered a portion of the sample to a 2 m Haysep-D column. The other sampling loop delivered the rest of the sample to a 1.82 m Molsieve-5 Å column with an attached 0.475 m Haysep-D pre-column. Each of these samples were directed to a pair of in-line detectors: a flame ionization detector (FID) for detecting carbon-containing products and a thermal conductivity detector (TCD) for other products. The use of a Molsieve-5 Å column necessitated the use of a timed injection-backflow sequence to prevent column fouling with CO₂. While this double injection scheme was capable of separating CO₂, CO, and O₂, it produced multiple signals for the same product, as well as introduced variation in the CO and O₂ intensity depending on the sample path.

The raw CO₂ conversion yield, X_{CO_2} (eqn (S1)†), was quantified by comparison to a measured calibration curve and used as the metric for catalytic activity. CO and O₂ retention times were calibrated *via* a commercial calibration standard. Prior to each plasma catalysis experiment, null runs were performed with the catalyst in place but without the plasma to establish the baseline quantity of CO₂ in the GC measurements (denoted CO_{2,in}). The plasma input power was then set to a determined value during each experiment and the output gas of the reactor was sampled to measure the amount of unreacted CO₂ in the exit stream (CO_{2,out}). The raw CO₂ conversion ratio, X_{CO_2} , was determined by the relative loss of CO₂ during the experiment after accounting for the gas expansion for this isobaric reaction.^{52–54} The details of this correction are shown in the ESI† (eqn (S1)–(S4)). This process was repeated three times on fresh, as-

prepared samples for each combination of catalyst material and reactor input power. The specific energy input (SEI) into the feedstock was defined by eqn (S3),† with the energy input varying by no more than 3% throughout the period of a single test. Each test consisted of a two-minute plasma-on period followed by a twenty-minute plasma-off and purge period. The isobaric-corrected CO₂ conversion percent values, (f_{CO_2}), and the SEI were utilized to determine the input power conversion of the process as in eqn (S4).† The gas residence time in the discharge region was assumed to be the same for all packed reactor conditions due to the highly uniform nature of the catalyst/support pair and the consistent sample-to-sample packing density.

Imaging and spectral characterization

Scanning electron microscopy (SEM) imaging was performed using a FEI Nova NanoLab 200 SEM equipped with a field emission source and an energy-dispersive X-ray (EDX) spectrometer. X-ray photoelectron spectroscopy (XPS) measurements were taken with PHI Versaprobe Tetra.

Optical microscopy was performed using an Olympus BX53M microscope under white light illumination. Thermal imaging was performed using a thermal camera (Shot-Pro, SEEK Thermal).

Results and discussion

Effect of ALD catalyst on CO₂ conversion and reactor input power conversion efficiency

We measured the conversion percentage of CO₂ resulting from a single pass through the discharge region of the DBD plasma reactor. The conversion ratio of each catalyst is shown in Fig. 2a as a function of input power. The conversion percentage of CO₂ increased significantly as a function of input power for each tested catalyst. The trend of increasing CO₂ conversion rate with increasing input power, regardless of catalyst, has been observed in previous work, and is typically attributed to both an increase in the frequency of micro-discharges and an increase in the average electron energy in the discharge gap.^{44,55,56} However, we observed a diminishing relationship between CO₂ conversion



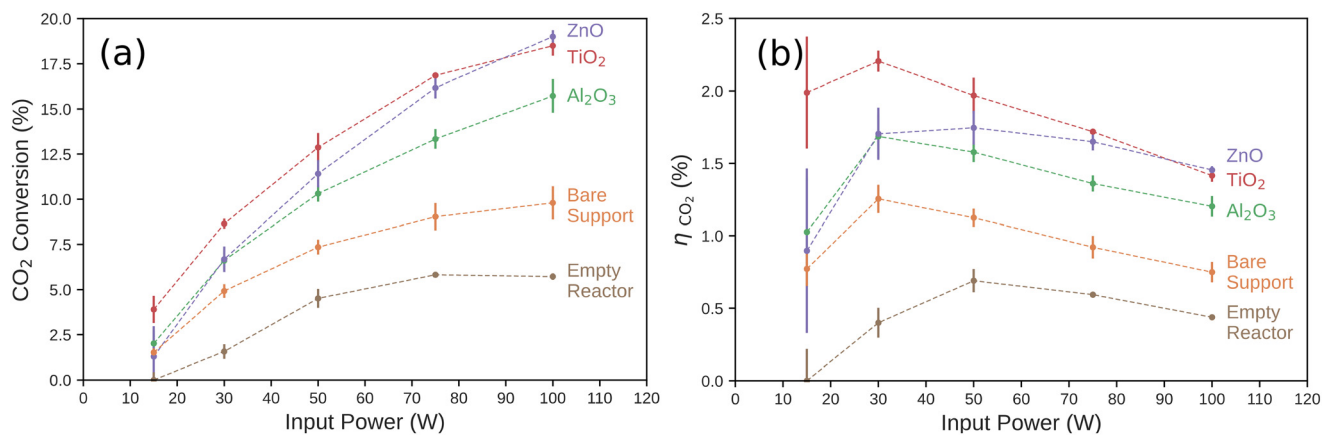


Fig. 2 (a) CO₂ conversion percentage and (b) the power conversion efficiency (η_{CO_2}) for each catalyst as a function of input power to the DBD plasma reactor. The conversion percentage and error bars are calculated from experiments performed on three independent, identically prepared catalysts and show the range of the data collected. η_{CO_2} was computed from the CO₂ conversion rate for each catalyst (Fig. 2a) and the Gibbs free energy of the CO oxidation reaction as described in eqn (S4).†

rate and input power of the reactor in all cases. The conversion ratio and input power conversion efficiency in the empty reactor, as shown in Fig. 2b, was consistently lower than the reactor loaded with the bare support material or any of the tested catalyst at all input powers. While it is not possible to exclude the heterogeneous effects of the quartz tube surface, we used the empty tube measurements as the baseline for the direct, uncatalyzed plasma reaction process. It is also not possible to isolate surface interactions from the mass transport effects of the bare support. The difference in reactivity between the empty reactor and the bare support while the surface composition is practically the same (significantly SiO₂) does indicate that there are non-catalytic effects of making the pathway through the reactor more tortuous. This is likely due to the characteristics of the plasma-support interactions and to the increased residence time of the gas in the plasma volume.

We then compared the performance of catalytic effects of ultra-thin ALD films deposited on the support material. While the Al₂O₃ catalyst was the least active of the tested layers, it was more active than the bare support, especially at higher input powers. The TiO₂ catalyst showed the highest conversion percentage across the input power range, with a 3.9% conversion rate at 15 W to an 18.5% conversion rate at 100 W. We observed a power dependence on the behavior of ZnO. At the lowest power measured (15 W), the ZnO catalyst performed worse than the bare support, though these were both within one standard deviation of each other and Al₂O₃. ZnO and Al₂O₃ were statistically equivalent at 30 W. At higher input power, the CO₂ conversion percentage rose steeply with power compared to all other catalysts measured. At 100 W, ZnO gave the highest CO₂ conversion percentage (19.0%) of any of the catalysts measured here. This is significant as ZnO has traditionally been utilized as a support for catalysts in both thermal- and non-thermal CO₂ hydrogenation plasma reactions. ZnO has generally been shown to have negligible catalytic effect when used in DBD systems alone at low input

powers. It is usually thought to contribute to these reactions by acting as a support for other metal catalysts and forming reduced ZnO_x species near the catalyst-support interface.^{49,50,57}

Another important metric for fuel-forming reactions, such as CO₂ reduction to CO, is the input power conversion efficiency, η_{CO_2} . This measures the ratio of the Gibbs free energy of combustion of the moles of CO₂ converted to CO to the reactor input power. This is particularly important if the products are intended to be used as a fuel, for example. Here, the oxidation of CO(g) to CO₂(g) is the reverse of the main reaction occurring in the plasma reactor. CO(g) formation has a Gibbs free energy of reaction of $\Delta G_{\text{rxn}}^0 = -257.2 \text{ kJ mol}^{-1}$, and this is the simplest reaction to consider as the chemical power output of the plasma reactor. Other reactions or methods for recouping the input power, such as use of the waste heat, are not considered here. As shown in Fig. 2b, the input power conversion efficiency of the CO₂ reduction reaction generally decreased as input power increased. We observed a nearly linear decrease for input power conversion efficiency on the TiO₂ catalyst with input power. Compared to the TiO₂ catalysts, the other catalysts did not have comparable CO₂ conversion ratios at low input power. The difference between the power conversion efficiency of the ZnO and TiO₂ became insignificant at higher input powers. Also, the input power conversion efficiency on Al₂O₃ and the bare support all slightly increased at low input powers before also diminishing at higher input powers. These results suggest that the optimal operating input power of the reactor is dependent on the choice of metal oxide plasma catalyst.

We used a power law analysis to characterize the catalytic behavior of each of the catalyst materials. This empirical analysis is analogous to Tafel slope analysis of the potential-dependent performance of electrocatalysts and more sophisticated methods in plasma catalysis.^{58,59} Fig. 3 shows the square of the fraction of CO₂ conversion as a function of input power. TiO₂, Al₂O₃, and the bare support all showed





Fig. 3 Squared yield of converted CO₂, $(f_{CO_2})^2$, of each catalyst at a fixed CO₂ flow rate of 75 sccm as a function of input power. $(f_{CO_2})^2$ was computed from eqn (S2).†

linear behavior in this plot, indicating a square-root dependence of CO₂ conversion yield on the input power. The slopes of these lines may have some physical or kinetic relevance, but we can simply use them as a metric for catalytic performance. At low power (≤ 30 W in these experiments), ZnO follows the same slope as Al₂O₃, a known support material for catalysts, indicating low catalytic activity. The slope increases significantly at higher powers, yet remains linear and intersects the TiO₂ line at 100 W. This suggests that the ZnO behavior is power dependent, switching from behaving as a support at low power to a catalyst at high powers. This may affect the interpretation of ZnO-supported catalytic behavior in other systems.

We used optical microscopy to characterize the surface of the catalysts before and after plasma exposure. This helped to determine if any structural changes occurred during the

experiments. As seen in Fig. 4, the support fibers appear unchanged after several minutes of plasma exposure, with no indication of melting or degradation. Small, optically opaque particles were observed on each of the catalysts after 30 minutes of plasma exposure. EDX spectroscopy of the particles and on regions away from the particles, Fig. 5, showed that the particles were composed of a significant amount of carbon. We attribute the Cu signal present in Fig. 5c to the conductive tape used to mount the sample to the stage. The substrate Cu signal was more intense due to the relatively lower electron-beam attenuation from the fiber compared to the spectrum measured in Fig. 5b. The EDX composite, Fig. 5b, also showed that the regions of the C deposition were mainly localized to the particles, and only adventitious carbon or very thin deposits were visible elsewhere on the fibers. The particles were approximately 1–3 μm in diameter and were observed more frequently at the intersections of support fibers or at irregularities in the fiber diameter. We observed that the amount of carbon deposition appeared to increase with increasing reactor input power. Qualitatively, greater deposition was visible *via* optical microscopy on catalysts exposed to 100 W of plasma input power than 50 W. The Al₂O₃ catalysts saw significantly more carbon deposits than either the TiO₂ or ZnO catalysts. In some cases, the Al₂O₃ catalysts were visually darker in some regions due to the particle deposition, whereas negligible discoloration was observed on the TiO₂- and ZnO-coated catalysts. Deposition of carbon species in DBD reactors due to the Boudouard reaction has been previously reported at the input powers and temperatures encountered during the experiments.^{19,60,61} Due to the gas-phase analysis used here, the direct splitting of CO₂ to produce C(s) cannot be quantified. However, these observations suggest that a portion of the CO₂ transformation activity on Al₂O₃ can be ascribed to coke formation. More generally, the deposition of carbon may vary with catalyst chemistry.¹² Additionally, the CO₂ transformation rates for each catalyst showed negligible differences as we repeated the power series measurements on a given catalyst. While some solid carbon could have deposited during this sequence, the negligible variation from the first power series measurement (fresh catalyst) to the third series (potential C deposition) suggests that any C(s) present does not influence the catalytic activity of the metal oxide layer.

XPS is sensitive to the composition and oxidation state of an interface. We used XPS to characterize the surface of ALD-synthesized catalysts in both the as-prepared and post-plasma exposure states. This comparison allowed us to measure any transformations or degradation of the catalyst during the plasma CO₂ reduction process. All *ex situ* samples were left in the air for at least 1 day prior to XPS characterization. The TiO₂ catalyst showed intensity at binding energies of 458.2 eV and 464 eV corresponding to the Ti 2p_{3/2} and the 2p_{1/2} features, respectively, consistent with the Ti⁴⁺ in TiO₂ (Fig. 6a).^{44,62} After plasma exposure, we observed the same features. Additionally, we observed the

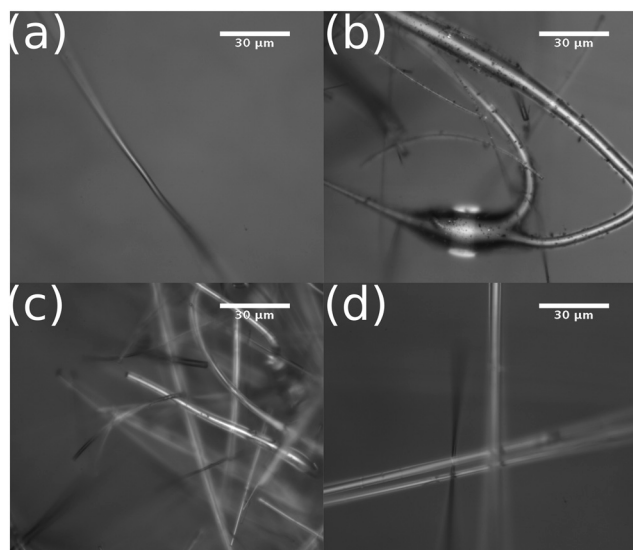


Fig. 4 Optical microscopy of (a) an ALD-TiO₂ coated support prior to plasma exposure, (b) an ALD-Al₂O₃ coated support after plasma catalysis, (c) an ALD-TiO₂ coated support after plasma catalysis, and (d) an ALD-ZnO coated support after plasma catalysis.





Fig. 5 (a) An overlaid composite image of the secondary electron micrograph (grayscale) and the C EDX micrograph (red) of a post-reaction Al₂O₃ plasma catalyst. EDX spectra of an Al₂O₃ plasma catalyst after the reaction (b) on the supported catalyst away from any deposited particle and (c) on a particle deposited during the DBD plasma reaction.

formation of a new feature in these XPS regions at binding energies of 456.2 eV and 461 eV, consistent with formation of reduced Ti³⁺ species as shown in other plasma-reduced TiO₂ surfaces, possibly as Ti₂O₃.^{54,63,64} A feature at higher binding energy, 532 eV, was also observed in the O 1s XPS region on

the TiO₂ catalyst before and after plasma exposure (Fig. 6b), however this feature is likely the result of water chemisorbed onto the surface of samples.⁶⁵ The lower binding energy feature in the same O 1s spectra is likely the result of hydroxide or other oxygen-containing adsorbates.⁶⁵ The integrated intensity of both features remained consistent before and after plasma exposure. Additionally, negligible intensity above background was observed for the Al 2p and Si 2p XPS regions that would indicate the displacement of the catalyst layer and exposure of the base layer (Fig. S4 and S5[†]). The integrated intensity of the Ti 2p_{3/2} XPS peak (including the Ti³⁺ feature) was preserved between the experiments. The underlying support was not observed in the post-plasma measurements. This indicates that the TiO₂ catalyst layer was unbreached during the plasma catalysis over the course of the experiments described here.

We observed that the integrated intensity of the signal attributed to Zn 2p_{3/2}, Fig. 7, did not change significantly. Additionally we observed negligible intensity in the Al 2p and Si 2p XPS regions that would indicate the exposure of the base layers due to the disruption of the conformal ZnO catalyst (Fig. S4 and S5[†]). We also observed an additional feature in the O 1s XPS region, indicating adsorbed hydroxide on the surface.⁶⁵ ALD-synthesized Al₂O₃ did not have a significant change during the plasma CO₂ reduction process in either the Al 2p or O 1s XPS regions (Fig. 8a and b). While the Al 2p peak could not be used to assess the stability of this layer, no significant intensity was observed in the Si 2p XPS region of the post-plasma exposure Al₂O₃ catalyst.

It has been proposed that the decrease in performance for CO₂ splitting into CO and O₂ in a packed DBD reactor, as compared to a DBD reactor with no packing, is due to the modification of the discharge mode *via* a decrease in the number of filamentary discharges per unit time in the discharge region.^{57,66} In our system, the increase in performance by the bare support over the unpacked reactor suggests an increase in the number of overall discharges. The

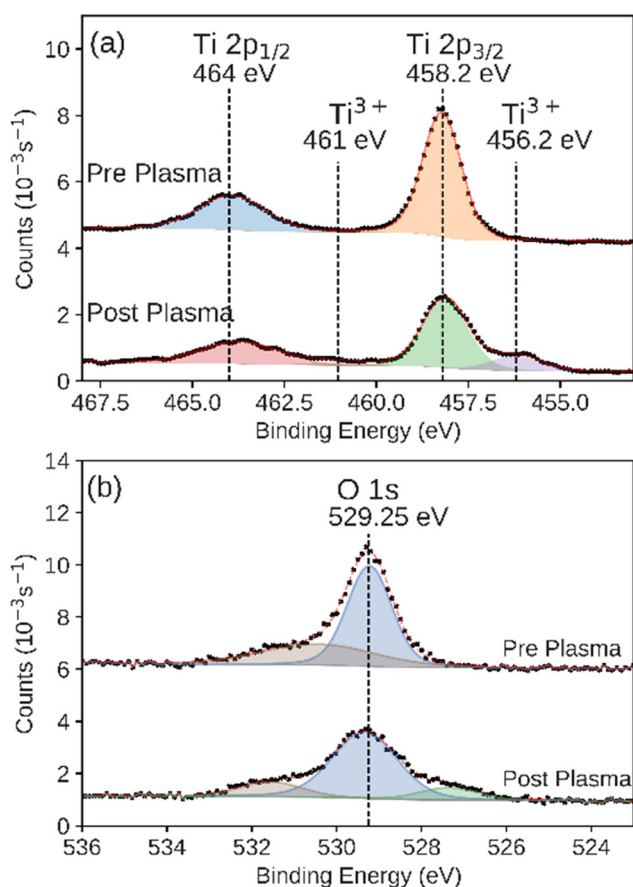


Fig. 6 XPS of the (a) Ti 2p and (b) O 1s regions measured on the as prepared and post-plasma ALD-synthesized TiO₂ catalyst. Black markers indicate raw data. Fitted peaks are shown in color with a red line showing the sum of the fit features.





Fig. 7 XPS of the (a) Zn 2p_{3/2} and (b) O 1s regions measured on the as-prepared and post-plasma ALD-synthesized ZnO catalyst. Black markers indicate raw data. Fitted peaks are shown in color with a red line showing the sum of the fit features.



Fig. 8 XPS of the (a) Al 2p and (b) O 1s regions measured on the as-prepared and post-plasma ALD-synthesized Al₂O₃ catalyst. Black markers indicate raw data. Fitted peaks are shown in color with a red line showing the sum of the fit features.

uniformity of the catalyst structure allowed us to regularize the dielectric material in the gap, and thus the overall nature of the plasma in the reactor. The exact growth rate of each material on the support is difficult to measure. We used rates determined by measuring deposition on planar Si wafers. While this is mostly the same growth mechanism, there is potentially some variability in the growth rates due to surface initiation and morphological differences. Thus, the catalyst layer thicknesses on the scaffold should only be considered accurate to one significant figure. The scaffold was therefore roughly 10⁴ thicker than any of the catalyst layers here. Any differences in CO₂ conversion can therefore be attributed to either the differences in catalytic activity for CO₂ reduction due to the ALD-synthesized catalyst layer or electric field enhancement effects due to the permittivity of the ultra-thin catalyst layers when exposed to plasma. The extremely thin catalyst layers should have a negligible effect on the physical characteristics of the plasma, and the dielectric properties of the reaction volume are dominated by the permittivity of the support material.^{67,68} Given this approximation, the differences in the plasma itself near the surface of the catalyst should be negligible. Therefore, this approach produces regularized plasma catalysts for which any

macroscopic changes in products can be ascribed to microscopic changes in activity. For example, the observed transition in the catalytic activity of ZnO is directly related to the reactivity of the material itself, rather than any morphological reconfiguration of the layer under plasma exposure. To our knowledge, this is the first report of this observation on ZnO.

Several factors can affect the interactions between CO₂ molecules and the catalyst surface, and therefore the catalytic activity. The catalysts studied here represent a range of Lewis basicity, with relative basicity increasing in the order TiO₂ < Al₂O₃ < ZnO.⁶⁹ It has been shown that a higher catalyst Lewis basicity promotes the chemisorption of CO₂ onto the surface, promoting CO₂ dissociation, however this trend was not observed in conversion studies.^{9,70} The formation of oxide vacancies (V_o) during plasma exposure tends to increase the basicity of the surface. These V_o have been shown to act as active sites for plasma CO₂ reduction and also increase the electron density of the adjacent metal cation and promote its interactions with the acidic carbon on CO₂.⁷¹⁻⁷⁴ The formation energies of V_o in Al₂O₃, TiO₂ and ZnO have been reported to be 7.0 eV, 5.7 eV, and 4.1 eV respectively.⁷¹ This suggests that the formation of V_o would be suppressed on



Al₂O₃ as compared to TiO₂ or ZnO due to the higher energies required. This is supported by the observation of the formation of reduced metal centers on the TiO₂ catalyst during CO₂ plasma exposure. While we observed the presence of reduced Ti³⁺ on the surface of the TiO₂ catalyst post plasma exposure, we did not observe reduced Zn in XPS measurements. We attribute this to the poor stability of V_o defects or reduced Zn in ZnO under air exposure.⁷⁵ The adsorption energy of CO₂ at an oxygen vacancy on the ZnO surface, has been calculated to be -0.60 eV.⁷⁶ This is less favorable than the adsorption of CO₂ onto an oxygen vacancy in TiO₂, whose energy ranges from -2.53 eV to -0.98 eV depending on the location of the vacancy and the orientation of CO₂.⁷⁴ The formation energy difference suggests that ZnO would be expected to have a greater number of oxygen vacancies present on its surface compared to TiO₂, particularly as the rate of the vacancy formation is increasing with input power. However, those vacancies would not be as capable of forming a bound CO₂ adduct intermediate as strongly as TiO₂. This might explain the slightly lower CO₂ conversion rate observed on ZnO compared to TiO₂ at low input power, where the populations of V_o would be relatively low and the greater CO₂ adsorption energy would result in greater CO₂ activation. At higher input powers, the greater population of available V_o in activated ZnO would contribute to the relatively improved performance over TiO₂.

The results described here reinforce the significance of the formation of V_o in DBD catalysts for CO₂ conversion. This also highlights the importance of understanding the contribution and activity of the support material when designing plasma CO₂ conversion systems, as support materials such as ALD-ZnO and ALD-Al₂O₃ have been shown to have small but significant power-dependent activities that can influence conversions beyond support-catalyst interactions and plasma enhancement effects. The ALD synthesis method described here can be used in conjunction with other techniques, such as surface specific spectroscopy and more advanced reactor output product analysis to characterize the contribution of a support material more fully. Care should be taken to not underestimate the activity of the support and experiments should be designed as to ensure the activity of the support controlled when comparing multiple catalysts across wide input power ranges.

Conclusions

We isolated the catalytic effects of metal oxide catalysts for DBD plasma CO₂ reduction processes by precisely synthesizing nanoscale layers with ALD. ALD allowed for control variations in structure, surface area, and electronic effects between catalysts. This allowed for the unambiguous study of the catalytic properties of each compound on under equivalent plasma and catalyst structure conditions. These studies show that these catalysts can withstand and be used for DBD plasma CO₂ reduction for meaningful lengths of time and can be deposited onto support structures such that

their affinities for CO₂ reduction can be studied while maintaining identical bulk reactor properties such as flow rate, packing density, surface area and surface structure. The highest conversion efficiency was 2.6% utilizing the ALD-TiO₂ at 15 W input power. ZnO, a common support material, showed an unexpected transition in its catalytic activity. At low power (15 W) it behaved like the bare support, while at high power (100 W) it yielded the highest overall CO₂ conversion percentage (19.0%) of any catalyst. Post-catalyst studies showed that the ALD TiO₂ and ZnO catalysts were resistant to carbon deposition when compared to ALD Al₂O₃. XPS studies showed the formation of reduced metal species on the TiO₂ surface after the plasma reaction, which suggests the formation of active oxygen vacancies. Decoupling the macrostructure of the catalyst from the catalytically active surface composition in this way removes potential ambiguities in the interpretation of differences in catalytic effects in plasma CO₂ reduction.

Data availability

Data will be made available by the authors on reasonable request.

Author contributions

Samuel K. Conlin: conceptualization, methodology, investigation, validation, writing – original draft. Hamed Mehrabi: methodology, validation, writing – review & editing. David N. Parette: methodology, validation, writing – review & editing. Eva M. Nichols: conceptualization, writing – review & editing, funding acquisition. Robert H. Coridan: conceptualization, validation, writing – original draft, supervision, funding acquisition.

Conflicts of interest

The authors declare no competing financial interest.

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