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Photocytotoxicity and photoinduced phosphine ligand exchange in a Ru(II) polypyridyl complex†

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Two new tris-heteroleptic Ru(II) complexes with triphenylphosphine (PPh₃) coordination, *cis*-[Ru(phen)₂(PPh₃)(CH₃CN)]²⁺ (**1a**, phen = 1,10-phenanthroline) and *cis*-[Ru(biq)(phen)(PPh₃)(CH₃CN)]²⁺ (**2a**, biq = 2,2'-biquinoline), were synthesized and characterized for photochemotherapeutic applications. Upon absorption of visible light, **1a** exchanges a CH₃CN ligand for a solvent water molecule. Surprisingly, the steady-state irradiation of **2a** followed by electronic absorption and NMR spectroscopies reveals the photosubstitution of the PPh₃ ligand. Phosphine photoinduced ligand exchange with visible light from a Ru(II) polypyridyl complex has not previously been reported, and calculations reveal that it results from a *trans*-type influence in the excited state. Complexes **1a** and **2a** are not toxic against the triple negative breast cancer cell line MDA-MB-231 in the dark, but upon irradiation with blue light, the activity of both complexes increases by factors of >4.2 and 5.8, respectively. Experiments with PPh₃ alone show that the phototoxicity observed for **2a** does not arise from the released phosphine ligand, indicating the role of the photochemically generated ruthenium aqua complex on the biological activity. These complexes represent a new design motif for the selective release of PPh₃ and CH₃CN for use in photochemotherapy.

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Introduction

Ruthenium(II) polypyridyl complexes exhibit useful excited state properties that have been explored in photochemotherapy (PCT), photodynamic therapy (PDT), and solar energy conversion, among other applications.^{1–8} The spatiotemporal control possible with these complexes shows promise in alternative cancer therapies, circumventing systemic toxicity present in traditional cancer therapies, such as approved platinum drugs.⁹ Typically, PCT and PDT agents are activated by the absorption of visible light in the irradiated area, leading to the population of excited states that can produce cytotoxic ¹O₂ for PDT or induce the release of a therapeutic agent in PCT. Unlike complexes used in photochemotherapy, PDT agents rely on the presence of oxygen, which can represent a drawback in the hypoxic environments found in solid tumors,^{10–12} making PCT agents an important area of research to advance photoinduced treatments.^{13–15}

Coordination to the Ru(II) center through a Lewis basic site, such as a nitrile or pyridine functional group, have been explored as PCT agents and for dual PCT/PDT activity, since many drugs that can be photoreleased possess one of these groups able to coordinate to a transition metal center.^{16–20} A frequent challenge, however, is the ability of other strong field ligands, such as phosphines, to undergo photoinduced dissociation. Importantly, molecules with a triphenylphosphonium group and cationic compounds with a triphenylphosphine (PPh₃) ligand have been shown to enhance cellular uptake,^{21,22} leading to an interest in the investigation of divalent ruthenium triphenylphosphine complexes for PCT.

Phosphine ligands, such as PMe₃ (Me = methyl) and PPh₃, have been shown to act as ancillary ligands that increase or promote the photosubstitution of other monodentate ligands in the Ru(II) coordination sphere.^{23–25} In particular, PR₃ (R = Me, Ph) ligands are generally stronger field ligands relative to N-coordinated pyridine and acetonitrile. Strong π-backbonding to phosphine ligands has also been used to modify the electronic structure on the ruthenium center to reduce the overpotential of CO₂ reduction catalysts and to tune the absorption and emission properties.^{26–29} Whereas CH₃CN and pyridine have been previously shown to undergo photoinduced ligand exchange in Ru(II) complexes, phosphine ligands are largely inert to photo-substitution.^{4,24,29–34} The design of complexes that can selectively photodissociate phosphines can enable the use of drugs with phosphine motifs in PCT,^{35–37} as well as the synthesis of supported catalysts patterned with selective irradiation.^{38–40}

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In the present work, two new heteroleptic Ru(II) complexes containing one PPh₃ and one CH₃CN ligand, *cis*-[Ru(phen)₂(PPh₃)(CH₃CN)]²⁺ (**1a**, phen = 1,10-phenanthroline) and *cis*-[Ru(biq)(phen)(PPh₃)(CH₃CN)]²⁺ (**2a**, biq = 2,2'-biquinoline), were synthesized and characterized, and their structures are shown in Fig. 1. The electronic absorption, electrochemistry, and photochemistry of **1a** and **2a** were investigated and compared to those of their bis-acetonitrile analogs, *cis*-[Ru(phen)₂(CH₃CN)₂]²⁺ (**1b**) and *cis*-[Ru(biq)(phen)(CH₃CN)₂]²⁺ (**2b**). Based on the steric distortion introduced by the bulky PPh₃ ligand, complexes **1a** and **2a** were expected to exhibit more facile CH₃CN dissociation. While **1a** exhibits photoinduced CH₃CN exchange upon visible light excitation, **2a** represents the first example of photoinduced exchange of a PPh₃ ligand from a Ru(II) polypyridyl complex, a surprising departure from the commonly observed substitutional inertness of PPh₃ ligands. Single-crystal X-ray structures of **1a**, **2a**, and the photoproduct of **2a** following photolysis in CH₃CN and pyridine (**1**), were collected and calculations were performed on **1a** and **2a** to gain better understanding of the origin of the unusual photoreactivity. In addition, complexes **1a** and **2a** were evaluated for their toxicity against the triple-negative breast cancer MDA-MB-231 cell line in the dark and upon irradiation. The present findings show enhanced activity following photoinduced ligand dissociation for both complexes and that PPh₃ release from **2a** results in a modest increase in toxicity as compared to CH₃CN photodissociation in **1a**. Importantly, both **1a** and **2a** exhibit significantly greater photoactivity than related complexes without PPh₃ in their coordination sphere. The present work is consistent with greater cellular uptake by the PPh₃-containing complexes, laying the groundwork for the design of new photoactive complexes with enhanced activity.

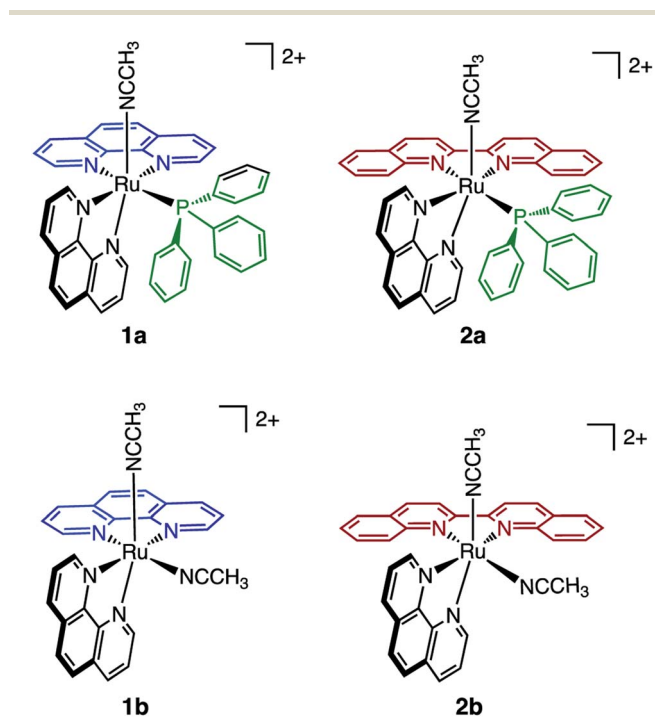


Fig. 1 Schematic representation of the molecular structures of **1a**, **1b**, **2a**, and **2b**.

Experimental

Materials

All materials were used as received without further purification, including 1,10-phenanthroline, 2,2'-biquinoline, CD₃CN, CD₃OD, (CD₃)₂CO, lithium chloride, pyridine, silver tetrafluoroborate, tetrabutylammonium hexafluorophosphate, and triphenylphosphine which were purchased from Sigma-Aldrich. Ethanol (200 proof) was obtained from Decon Laboratories, acetone, acetonitrile, dichloromethane, diethyl ether, *N,N*-dimethylformamide, 85% H₃PO₄, and toluene were acquired from Fischer Scientific, and ammonium hexafluorophosphate was purchased from Oakwood Chemical. Complexes **1b** and **2b**,⁴¹ [Ru(phen)₂Cl₂],⁴² [Ru(*p*-cymene)Cl₂],⁴³ and triphenylphosphine oxide⁴⁴ were prepared according to literature procedures.

[Ru(phen)₂(PPh₃)(Cl)](PF₆)

[Ru(phen)₂Cl₂] (0.16 g, 0.30 mmol), triphenylphosphine (0.14 g, 0.53 mmol), and excess LiCl were added to 10 mL ethanol/water (1 : 1, v/v) mixture sparged for 15 min with N₂. The reaction mixture was refluxed for 4 h under a nitrogen atmosphere, allowed to cool, concentrated by rotary evaporation, and then precipitated by adding it dropwise to a concentrated NH₄PF₆ solution. The product was purified by column chromatography, using a deactivated neutral alumina stationary phase and a 1 : 2 toluene : acetone mobile phase. The solvent was removed from the fraction containing the product *via* rotary evaporation, producing a dark orange solid (0.078 g, 29% yield). ¹H NMR (400 MHz, (CD₃)₂CO, Fig. S1†): δ 9.72 (d, 1H, *J* = 5.4 Hz), 8.77 (m, 2H), 8.68 (d, 1H, *J* = 8.2 Hz), 8.55 (dd, 1H, *J* = 5.7, 3.9 Hz), 8.37 (s, 2H), 8.36 (d, 1H, *J* = 1.3 Hz), 8.30 (d, 1H, *J* = 8.8), 8.23 (dd, 1H, *J* = 20.3, 9.0 Hz), 8.13 (d, 1H, *J* = 8.8 Hz), 7.90 (dd, 1H, *J* = 8.2, 5.3 Hz), 7.86 (dd, 1H, *J* = 8.2, 5.3 Hz), 7.77 (dd, 2H, *J* = 8.3, 5.3 Hz), 7.65 (d, 1H, 5.3 Hz), 7.53 (m, 2H), 7.37 (t, 5H, *J* = 8.7 Hz), 7.27 (t, 2H, *J* = 7.6 Hz), 7.11 (m, 6H). ³¹P{¹H} NMR (400 MHz, (CD₃)₂CO, Fig. S2†): δ 45.1 (s, 1P).

[Ru(phen)₂(PPh₃)(CH₃CN)](PF₆)₂ (**1a**)

[Ru(phen)₂(PPh₃Cl)](PF₆) (0.058 g, 0.064 mmol) was dissolved in 10 mL of acetonitrile/H₂O (1 : 1, v/v) mixture and, under an atmosphere of nitrogen, was refluxed overnight. After cooling to room temperature, the reaction solution was added dropwise to a concentrated aqueous NH₄PF₆ solution. The precipitate that formed was collected by filtering over Celite and purified on a neutral alumina column eluted with a 1 : 2 toluene : acetone mobile phase. The purified solution was collected and the solvent was removed *via* rotary evaporation, affording the desired product as a yellow-orange solid (0.031 g, 53% yield). ¹H NMR (400 MHz, CD₃CN, Fig. S3†): δ 9.41 (d, 1H, *J* = 5.3 Hz), 9.09 (d, 1H, *J* = 5.2 Hz), 8.82 (dd, 1H, *J* = 8.4, 1.2 Hz), 8.59 (dd, 1H, *J* = 8.3, 1.2 Hz), 8.55 (dd, 1H, *J* = 8.2, 1.3 Hz), 8.37 (dd, 1H, *J* = 8.3, 1.2 Hz), 8.24 (dd, 2H, *J* = 28, 8.8 Hz), 8.13 (dd, 2H, *J* = 18, 8.8 Hz), 7.90 (dd, 1H, *J* = 3.0, 5.3 Hz), 7.75 (dd, 1H, *J* = 3.0, 5.3 Hz), 7.49 (m, 1H), 7.43 (m, 1H), 7.37 (m, 4H), 7.19 (m, 7H), 7.03 (td, 6H, *J* = 9.4, 1.1 Hz), 2.17 (s, 3H). ³¹P{¹H} NMR (400 MHz, CD₃CN,



Fig. S4†): δ 45.3 (s, 1P). ESI-MS(+): $[\text{M-PF}_6]^+$ $m/z = 910.193$ (calc. $m/z = 910.123$).

[Ru(*p*-cymene)(phen)Cl]

[Ru(*p*-cymene)Cl₂]₂ (0.30 g, 0.50 mmol) and 1,10-phenanthroline (0.19 g, 1.1 mmol) were dissolved in 4 mL acetonitrile and refluxed under a nitrogen atmosphere for 2 h, during which time a change from a red to orange solution was observed. A yellow-orange solid was collected by filtering over Celite (0.37 g, 82% yield). ¹H NMR (400 MHz, CD₃OD, Fig. S5†): δ 9.84 (d, 2H, $J = 5.5$ Hz), 8.84 (d, 2H, $J = 8.3$ Hz), 8.21 (s, 2H), 8.11 (dd, 2H, $J = 8.3, 5.3$ Hz), 6.23 (d, 2H, $J = 6.5$ Hz), 6.00 (d, 2H, $J = 6.2$ Hz), 2.66 (q, 1H, $J = 7.0$ Hz), 2.27 (s, 3H), 0.99 (d, 6H, $J = 6.9$ Hz).

[Ru(biq)(phen)Cl₂]

[Ru(*p*-cymene)(phen)Cl] (0.33 g, 0.68 mmol), 2,2'-biquinoline (0.18 g, 0.69 mmol), and excess LiCl were dissolved in 2 mL *N,N*-dimethylformamide and refluxed under a nitrogen atmosphere for 90 min. After refluxing was complete, the reaction mixture was allowed to cool to room temperature and then added dropwise to 30 mL aqueous LiCl solution, producing a dark green solution. A dark green solid was collected *via* vacuum filtration and was rinsed three times each with 20 mL H₂O and 20 mL diethyl ether. The solid was dissolved using 1 L of a CH₂Cl₂/methanol (1 : 1, v/v) solvent mixture, which was then removed by rotary evaporation to afford the desired product as a dark green solid (0.18 g, 44% yield).

[Ru(biq)(phen)(PPh₃Cl)](PF₆)

[Ru(biq)(phen)Cl₂] (0.048 g, 0.080 mmol), triphenylphosphine (0.039 g, 0.15 mmol), and excess LiCl were added to a 10 mL ethanol/water (1 : 1, v/v) mixture sparged for 15 min with N₂. The reaction mixture was refluxed for 4 h under a nitrogen atmosphere, allowed to cool, concentrated by rotary evaporation, and then added dropwise to a concentrated NH₄PF₆ solution to produce a purple precipitate. The product was purified by column chromatography, using a neutral alumina stationary phase and an acetone mobile phase. The solvent was removed from the fraction containing the product *via* rotary evaporation producing a red-purple solid (0.052 g, 66% yield). ¹H NMR (400 MHz, (CD₃)₂CO, Fig. S6†): δ 10.21 (s, 1H), 8.90 (m, 4H), 8.61 (d, 2H, $J = 7.8$ Hz), 8.53 (d, 1H, $J = 8.1$ Hz), 8.30 (t, 2H, $J = 7.3$ Hz), 8.12 (d, 1H, $J = 8.2$ Hz), 8.06 (dd, 1H, $J = 8.1, 5.5$ Hz), 7.85 (dd, 3H, $J = 32.8, 9.0$ Hz), 7.63 (m, 3H), 7.52 (t, 1H, $J = 7.6$ Hz), 7.28 (t, 2H, 7.5 Hz), 7.17 (s, 3H), 6.95 (m, 12H). ³¹P{H} NMR (400 MHz, (CD₃)₂CO, Fig. S7†): δ 42.5 (s, 1P).

[Ru(biq)(phen)(PPh₃)(CH₃CN)](PF₆)₂ (2a)

[Ru(biq)(phen)(PPh₃Cl)](PF₆) (0.044 g, 0.038 mmol) and AgBF₄ (0.019 g, 0.099 mmol) were dissolved in 10 mL of a CH₃CN/H₂O (1 : 1, v/v) mixture and were refluxed overnight under a N₂ atmosphere. After cooling to room temperature, the reaction mixture was added dropwise to a concentrated NH₄PF₆ solution, precipitate was collected by filtering over Celite, and then purified *via* bulk recrystallization using vapor diffusion of ether

into a concentrated solution of 2a in acetonitrile, which afforded a red-orange solid (0.013 g, 35% yield). ¹H NMR (400 MHz, CD₃CN, Fig. S8†): δ 10.1 (d, 1H, $J = 5.3$ Hz), 8.96 (d, 2H, $J = 8.7$ Hz), 8.69 (d, 2H, $J = 8.8$ Hz), 8.30 (m, 5H), 8.19 (m, 5H), 8.07 (dd, 2H, $J = 5.4, 2.8$ Hz), 7.64 (m, 4H), 7.54 (m, 5H), 7.28 (td, 3H, $J = 7.4, 1.1$ Hz), 7.21 (td, 3H, $J = 7.8, 1.5$ Hz), 7.16 (td, 3H, $J = 7.9, 1.4$ Hz), 2.34 (s, 3H). ³¹P{H} NMR (400 MHz, CD₃CN, Fig. S9†): δ 42.3 (s, 1P). ESI-MS(+): $[\text{M-PF}_6]^+$ $m/z = 986.252$ (calc. $m/z = 986.155$).

Instrumentation and methods

Electronic absorption spectra were collected using a Hewlett-Packard 8454 diode array spectrophotometer in 1 cm × 1 cm quartz cuvettes. The irradiation source for photolysis experiments was a 150 W Xe arc lamp (UHSIO) in a MilliArc lamp housing unit equipped with an LPS-220 power supply and an LPS-221 igniter (PTI). Irradiation wavelengths for quantum yield determination were selected by using bandpass and long-pass filters (CVI Melles Griot). Samples for photolysis were prepared under red light, sealed in an NMR tube (NMR) or cuvette (UV-Vis), and purged with N₂ for 15 min prior to irradiation.

The ¹H and ³¹P{H} NMR spectra were obtained using a Bruker 400 MHz DPX instrument in CD₃CN, (CD₃)₂CO, or CD₃OD. ¹H chemical shifts were referenced to the residual protonated solvent peak and ³¹P{H} shifts were referenced to an external 85% H₃PO₄ standard (0 ppm). ¹H and ³¹P{H} NMR photolysis experiments were performed in CD₃CN. Electrospray ionization mass spectrometry (ESI-MS) was performed using a Bruker microTOF instrument. For ESI-MS experiments, samples were dissolved in CH₃CN and referenced to a sodium trifluoroacetate standard.

Electrochemistry experiments were performed on a BASI model CV-50 W voltammetric analyzer (Bioanalytical Systems, Inc.) with a three-electrode cell utilizing a glassy carbon working electrode, a Pt wire auxiliary electrode, and a saturated Ag/AgCl (3 M NaCl) reference electrode. Samples were dissolved in acetonitrile containing 0.1 M tetrabutylammonium hexafluorophosphate (TBAPF₆) as an electrolyte, data was collected at a scan rate of 200 mV s⁻¹, and ferrocene was added at the end of each experiment as an internal reference (+0.43 V vs. Ag/AgCl in acetonitrile).⁴⁵

Crystals suitable for X-ray diffraction were obtained through vapor diffusion of diethyl ether into concentrated acetonitrile or pyridine solutions of the desired complex. Single crystal X-ray diffraction for 2a was performed using a dark red rectangular plate crystal in a nitrogen gas stream at 150 K. The diffraction pattern was collected using a Nonius Kappa APEXII CCD diffractometer and Mo K_α radiation ($\lambda = 0.7107 \text{ \AA}$). Data were integrated using the Bruker SAINT software program and scaled using the SADABS software program. Structures were solved and refined with the Bruker SHELXT Software Package within APEX2 and Olex2. Other single crystal X-ray diffraction measurements were performed on a Bruker Kappa Photon II CPAD diffractometer equipped with Mo K_α radiation (1a and 2b, $\lambda = 0.71073 \text{ \AA}$) or Cu K_α radiation (I, $\lambda = 1.54178 \text{ \AA}$) using a dark red crystal (1a), orange plate (2b), or red blade (I) in a nitrogen gas stream



at 100(2) K. The data were integrated using the Bruker SAINT software program and scaled using the SADABS software program. Solution by direct methods (SHELXT) produced a complete phasing model consistent with the proposed structure.

Spin restricted and unrestricted density functional theory (DFT) calculations were performed using the Gaussian09 program package.⁴⁶ Geometry optimizations and vibrational frequency calculations were performed with the SDD⁴⁷ basis set on Ru and the TZVP⁴⁸ basis set on all other atoms with the PBE exchange–correlation functional.^{49,50} The geometries of **1a**, **2a**, and **2b** were fully optimized starting from X-ray crystal structures and were verified to have positive harmonic frequencies, confirming the calculated structures as electronic energy minima. Molecular orbital calculations utilized the hybrid functional B3LYP,^{51–53} with the SDD basis set on Ru and the TZVP basis set on all other atoms. Spin densities were calculated using Mulliken population analysis (MPA) methods. Molecular orbitals from the Gaussian calculations were plotted using the Chemcraft program,⁵⁴ and the analysis of the molecular orbitals and Mayer bond order calculations were performed using AOMix-FO within the AOMix program.^{55,56}

The cell viability of all the synthesized complexes were determined by plating MDA-MB-231 cells in a 96 well plate at a density of 7000 cells per well in Dulbecco's Modified Eagle's Medium (DMEM) supplemented with 10% FBS and 1000 units per mL penicillin/streptomycin. The plates were incubated overnight in a 37 °C humidified incubator ventilated with 5% CO₂. The media was aspirated off and then quadruplicate wells were treated with DMEM supplemented with 10% FBS and 1000 units per mL penicillin/streptomycin containing different concentrations (30 μM to 500 nM) of the synthesized complexes in 1% DMSO. Plates containing wells with no cells were designated as blank wells whereas wells with cells that were not treated with the compound but only DMEM supplemented with 10% FBS and 1000 units per mL penicillin/streptomycin containing 1% DMSO (vehicle) were designated as control wells. The plates were then again incubated in a 37 °C humidified incubator ventilated with 5% CO₂. After 1 h of incubation, the cells were either irradiated with blue light ($t_{\text{irr}} = 20$ min, $\lambda_{\text{irr}} = 460\text{--}470$ nm, 56 J cm⁻²) or kept in the dark. After 20 minutes, the plates were placed in a 37 °C humidified incubator with 5% CO₂ for 72 h, after which time, 10 μL of 3-(4,5-dimethylthiazol-2-yl)diphenyl tetrazolium bromide (MTT) reagent (5 mg mL⁻¹ in PBS) was added to each well of the 96 well plate and incubated in a 37 °C humidified incubator ventilated with 5% CO₂ for 2 h. The media was then aspirated off and 100 μL of DMSO was added. The plates were then shaken for 20 min to ensure complete dissolution of the purple formazan crystals formed. Absorbance of each well was then measured at 570 nm. The mean absorbance values of the blank wells were calculated and subtracted from absorbance values for each well treated with a certain concentration of a compound. The absorbance of the control wells was also taken and subtracted with the average of the blank wells. The mean of these corrected control absorbances were then calculated. Viability of the cells was finally determined by dividing the corrected absorbance of the

compound wells by the mean corrected absorbance of the control wells and expressing the mean of the ratio as a percentage value. The % viability was plotted against the log of concentration (in molarity) of the compounds and the antilog of the concentration value at 50% viability was used to determine the EC₅₀ value of each complex against MDA-MB-231 cells.

Results and discussion

Electronic absorption and electrochemistry

The electronic absorption spectra of **1a**, **1b**, **2a**, and **2b** are shown in Fig. 2 and the corresponding absorption maxima and extinction coefficients are listed in Table 1. The singlet metal-to-ligand charge transfer (¹MLCT) absorption maxima of **1a** in CH₃CN, attributed to Ru(dπ)→phen(π*) transitions, are observed at 372 nm and 411 nm. For **2a** the Ru(dπ)→phen(π*) and Ru(ππ)→biq(π*) ¹MLCT bands are observed at 407 nm and 477 nm, respectively, in CH₃CN. The substitution of one of the phen ligands in **1a** for biq in **2a** results in a bathochromic shift in the ¹MLCT absorption maximum, as expected from the increased conjugation and subsequently increased π-accepting character of the biq ligand as compared to phen. A similar shift is observed in the ¹MLCT maxima of the bis-acetonitrile analogs, **1b** and **2b**, at 420 nm and 497 nm, respectively. In addition, the lower energy of the ¹MLCT transitions in **1b** and **2b**, as compared to the corresponding peaks in **1a** and **2a**, are consistent with the increased π-accepting character of the phosphine ligand as compared to acetonitrile.^{26,31,57–59}

The electrochemical reduction potentials for complexes **1a** and **2a** obtained from cyclic voltammetry (CV) experiments are listed in Table 1 and the corresponding CVs are shown in Fig. S10,† and are compared to those previously reported for **1b** and **2b**.⁴¹ The first reversible reduction events of **1a** and **1b** are localized on one of the phen ligands, with $E_{1/2}$ values at -1.29 V and -1.34 V vs. Ag/AgCl, respectively, and compare well to those reported for related complexes.^{60,61} In contrast, the first reduction couples of **2a** and **2b** observed at -0.82 V and -0.86 V vs. Ag/AgCl, respectively, are centered on the biquinoline ligand in each complex, consistent with the lower energy lowest unoccupied π* orbital in biq and similar to those measured in related complexes.^{60,62}

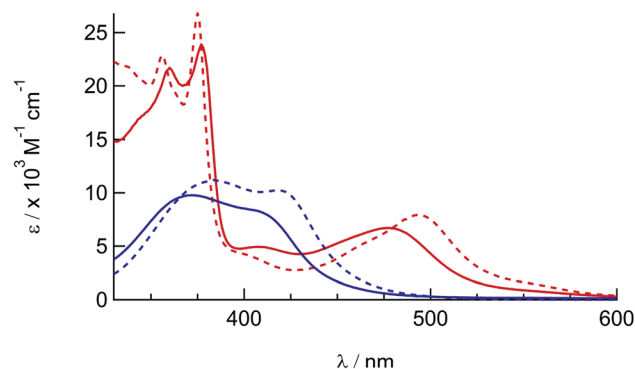


Fig. 2 Electronic absorption spectra of **1a** (solid blue), **1b** (dashed blue), **2a** (solid red), and **2b** (dashed red) in CH₃CN.



Table 1 Electronic absorption maxima (λ_{abs}), molar absorption coefficients (ϵ), and electrochemical half-wave reduction potentials ($E_{1/2}$) in CH_3CN

Complex	$\lambda_{\text{abs}}/\text{nm}$ ($\epsilon/\times 10^3 \text{ M}^{-1} \text{ cm}^{-1}$)	$E_{1/2}/\text{V}^a$
1a	372 (9.1), 411 (7.4)	+1.62, -1.29, -1.51
1b^b	383 (11.2), 420 (10.2)	+1.50, -1.34, -1.50
2a	360 (18), 377 (20), 407 (4.2), 477 (5.7)	+1.67, -0.82, -1.36 ^c
2b^b	356 (23), 375 (27), 406 (3.9), 497 (7.8)	+1.55, -0.86, -1.40

^a 0.1 M TBAPF₆, vs. Ag/AgCl in CH_3CN . ^b From ref. 41. ^c Irreversible.

The second reduction wave is localized on the phen ligand in **2a** and **2b**, observed at -1.36 V and -1.40 V vs. Ag/AgCl, respectively, and on the remaining phen ligand in **1a** and **1b**, at -1.51 V and -1.50 V vs. Ag/AgCl, respectively. The reversible oxidation events ranging from +1.50 to +1.67 V vs. Ag/AgCl are assigned to the Ru^{III/II} redox couple (Fig. S10† and Table 1). The ~120 mV shift of the Ru^{III/II} couples to more positive potentials in **1a** and **2a** relative to those in **1b** and **2b**, respectively, is consistent with the greater π -accepting character of triphenylphosphine, stabilizing the highest occupied molecular orbital (HOMO). Taken together, the electrochemical data indicate that the synthetic substitution of PPh₃ in **1a** and **2a** for CH₃CN in **1b** and **2b** primarily affects the energy of the Ru(d π) t_{2g}-type orbitals.

Photochemistry

In order to explore the light-induced ligand dissociation in complexes **1a** and **2a**, their photoreactivity was investigated by monitoring changes in the electronic absorption and ¹H and ³¹P {H} NMR spectra as a function of irradiation time. Irradiation of **1a** in water (<5% acetone) with visible light results in a decrease in intensity of the absorption peak at 372 nm and a concomitant increase in the 430–550 nm range with a shoulder at 455 nm, along with an isosbestic point at 407 nm (Fig. 3a). The presence of the isosbestic point is indicative of the reaction proceeding from the starting material to a single product. The changes to the ¹H NMR spectrum of **1a** in CD₃CN were also monitored as a function of irradiation time, resulting in a decrease in the resonance at 2.17 ppm associated with CH₃CN bound to ruthenium and the concomitant appearance of a resonance at 1.96 ppm, corresponding to free CH₃CN (Fig. 3b). These data indicate that the irradiation of **1a** results in the substitution of the CH₃CN ligand with a solvent molecule, in this case CD₃CN, with the absence of any additional photochemical reactions. The bathochromic shift in Fig. 3a is also consistent with this conclusion, as the photolysis of **1a** in water results in the formation of *cis*-[Ru(phen)₂(PPh₃)(H₂O)]²⁺, where the bound CH₃CN is substituted for the weaker-field, π -donating H₂O ligand, thus raising the energy of the Ru(dp) t_{2g}-type set and lowering the energy of the ¹MLCT transition.^{34,63,64}

The irradiation of **2a** in CH₃CN results in a decrease in the absorption at 407 nm and an increase a peak at 497 nm, with two isosbestic points at 396 nm and 478 nm (Fig. 4a). As shown in Fig. 4b, the spectrum of the photoproduct is nearly identical to that of **2b**, providing evidence that the irradiation of **2a**

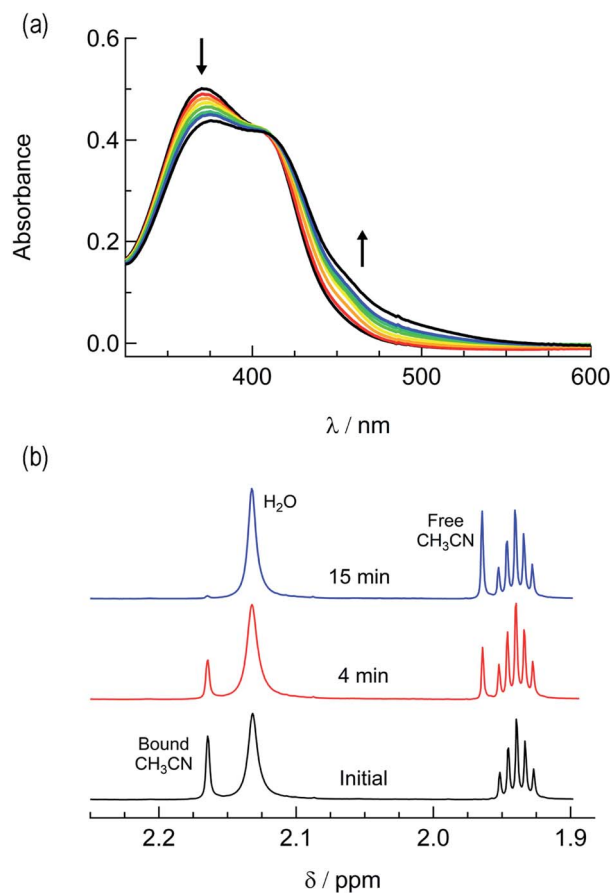


Fig. 3 Changes following the irradiation of **1a** ($\lambda_{\text{irr}} \geq 395 \text{ nm}$) to the (a) electronic absorption spectrum in H_2O , $t_{\text{irr}} = 0\text{--}30 \text{ min}$, and (b) ¹H NMR spectrum in CD_3CN , $t_{\text{irr}} = 0, 4, \text{ and } 14 \text{ min}$.

results in the photoinduced dissociation of the PPh₃ ligand generating *cis*-[Ru(biq)(phen)(CH₃CN)₂]²⁺, compound **2b**. The changes in the ¹H NMR spectra of **2a** in CD₃CN recorded as a function of irradiation time are also consistent with the exchange of the phosphine ligand following visible light irradiation (Fig. 5a). For example, the resonance at 2.34 ppm, associated with the ruthenium-bound CH₃CN ligand, decreases in intensity upon irradiation, with the concomitant growth of a peak at 2.47 ppm, associated with the photoproduct *cis*-[Ru(biq)(phen)(CH₃CN)(CD₃CN)]²⁺, similar to the resonances observed for the coordinated CH₃CN ligands in **2b**.⁴¹

The changes to the ³¹P{H} NMR spectra of **2a** upon irradiation provide additional evidence for PPh₃ exchange, where a decrease in intensity of the ³¹P{H} resonance at 42.3 ppm, associated with coordinated PPh₃ is observed during the photolysis (Fig. 5b). Concurrently, a ³¹P{H} resonance corresponding to triphenylphosphine oxide centered at 26.0 ppm appeared as a function of irradiation time (Fig. 4d and S11†). Importantly, following 15 minutes of irradiation of **2a**, the ³¹P{H} resonance associated with bound PPh₃ completely disappears, while the ¹H peaks of the photoproduct *cis*-[Ru(biq)(phen)(CH₃CN)(CD₃CN)]²⁺ persisted. In addition, the resonance at 1.96 ppm associated with free CH₃CN appears



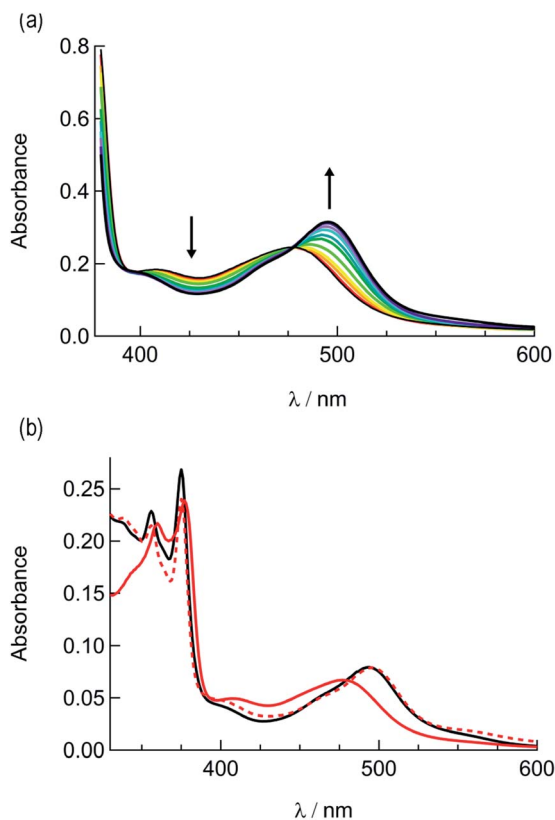


Fig. 4 (a) Changes in the electronic absorption spectrum of **2a** in CH_3CN following irradiation, $t_{\text{irr}} = 0\text{--}5$ min and (b) electronic absorption spectra of **2a** before irradiation (solid red), following 30 min irradiation (dashed red), and **2b** (black).

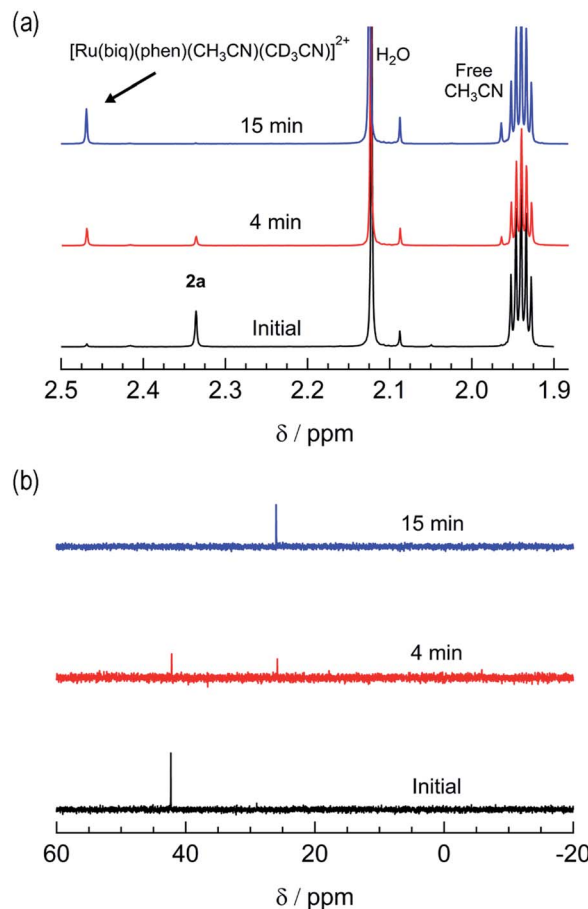


Fig. 5 Changes in the (a) ^1H and (b) $^{31}\text{P}\{\text{H}\}$ NMR spectra of **2a** in CD_3CN at $t_{\text{irr}} = 0, 4,$ and 15 min ($\lambda_{\text{irr}} \geq 395$ nm).

concomitantly with the peak at 2.47 ppm, indicating CH_3CN is also photodissociated albeit not completely on the timescale of the NMR photolysis experiment.

Based on these results, the question remains whether the irradiation of **2a** results in dissociation of both PPh_3 and CH_3CN from the starting material, or if CH_3CN exchanges only from the intermediate photoproduct $\text{cis}[\text{Ru}(\text{biq})(\text{phen})(\text{CH}_3\text{CN})(\text{CD}_3\text{CN})]^{2+}$ after the initial dissociation of PPh_3 from the starting complex. In an effort to address this point and trap the product of the first ligand exchange step, photolysis experiments were performed in the coordinating solvent pyridine under identical illumination conditions to those previously discussed in acetonitrile, and the results are shown in Fig. 6. Inspection of Fig. 6a reveals one set of isosbestic points at early irradiation times, up to ~ 2.5 min, observed at 316 nm, 448 nm, and 500 nm, and the decrease of the peak associated with **2a** at 490 nm with the appearance of a band at 540 nm. A second set of isosbestic points is evident at later times (Fig. 6b), from ~ 3.5 min to 20 min, at 338 nm, 361 nm, and 550 nm, with a loss of the species with absorption at 540 nm and the formation of the final product with maximum at 590 nm. These results point at the formation of solely one initial intermediate, **I**, with maximum at 540 nm, which then goes on to exchange a second ligand to generate the final product with a peak at 590 nm, assigned to $\text{cis}[\text{Ru}(\text{biq})(\text{phen})(\text{py})_2]^{2+}$ (**3**). As expected from the ability of CH_3CN to π -backbond with the $\text{Ru}(\text{d}\pi) t_{2g}$ -type

orbitals that is not present in pyridine, the $^1\text{MLCT}$ maximum of $\text{cis}[\text{Ru}(\text{biq})(\text{phen})(\text{py})_2]^{2+}$, **3**, is red-shifted compared to that of the product **2b**, $\text{cis}[\text{Ru}(\text{biq})(\text{phen})(\text{CH}_3\text{CN})_2]^{2+}$, in Fig. 4a.

The intermediate **I** was identified as $\text{cis}[\text{Ru}(\text{biq})(\text{phen})(\text{py})(\text{CH}_3\text{CN})]^{2+}$, generated by absorption of a single photon by **2a** and photosubstitution of the triphenylphosphine ligand for a solvent pyridine molecule. Identification was supported, in conjunction with $^{31}\text{P}\{\text{H}\}$ NMR spectra (Fig. 5b), by obtaining a single-crystal X-ray diffraction structure of the photoproduct generated by irradiating a solution of **2a** in pyridine with ≥ 395 nm light for 90 s (Fig. S12 and Table S1†). No further spectral changes took place after the conclusion of irradiation, indicating **I** is stable in the dark. This conclusion further supported by the persistence of the complex in a pyridine solution as the compound recrystallized *via* diethyl ether diffusion.

The presence of the two sets of isosbestic points permits determination of time-dependent concentrations of the three individual species in solution: **2a**, **I**, and **3**, during the photolysis, where **I** represents an intermediate species. The deconvolution of the associated spectra is possible from the known absorption spectra and molar extinction coefficients of **2a** and **3**, and the details of the analysis are presented in the ESI, Table



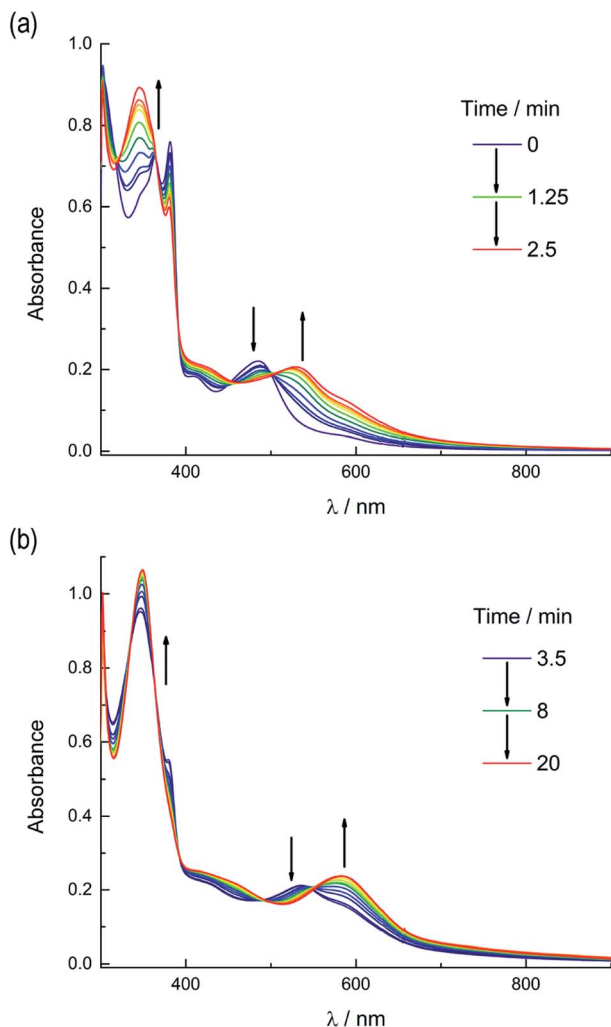


Fig. 6 Changes following irradiation ($\lambda_{\text{irr}} \geq 395$ nm) in the electronic absorption spectrum of **2a** in pyridine from (a) $t_{\text{irr}} = 0$ –2.5 min and (b) $t_{\text{irr}} = 3.5$ –20 min.

S2, and Fig. S13.† The spectra of **2a** and **3**, along with that of the intermediate, **I**, are shown in Fig. S13,† and the time dependent mole fractions of each species were calculated over the course of the photolysis and are displayed in Fig. 7. Fig. 7 shows that the loss of **2a** occurs rapidly, whereby at 70 s, 50% of the starting material remains and no amount is appreciable beyond 300 s. The formation of the intermediate **I** begins as early as 50 s of irradiation and reaches 99% conversion to the final product, **3**, at ~ 900 s. A maximum fraction of $\sim 55\%$ of the intermediate **I** is apparent at ~ 150 s (Fig. 7). From the known proportions of the three species the extinction coefficient for the intermediate was estimated and compared to initial and final product values in Fig. S13.†

The PPh_3 ligand photodissociation from **2a** apparent from the sequential formation of photoproducts in Fig. 6 and 7, as well as in the crystal structure of **I**, shows that irradiation of **2a** does not result in the photodissociation of the CH_3CN ligand, such that there is no evidence of the formation of $[\text{Ru}(\text{biq})(\text{phen})(\text{PPh}_3)(\text{py})]^{2+}$ following irradiation. Instead

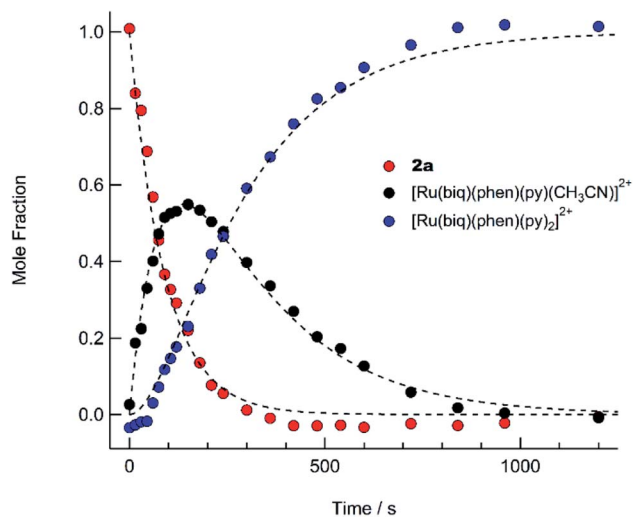


Fig. 7 Time dependent concentrations of **2a** (red circles), $\text{cis-}[\text{Ru}(\text{biq})(\text{phen})(\text{py})(\text{CH}_3\text{CN})]^{2+}$ (**I**, black circles), and $[\text{Ru}(\text{biq})(\text{phen})(\text{py})_2]^{2+}$ following irradiation ($\lambda_{\text{irr}} \geq 395$ nm) in pyridine from $t = 0$ to $t = 1200$ s. The dashed lines are least-squares fits to a consecutive reaction model with a system of equations describing the time-dependent concentration of each compound (see text).

CH_3CN substitution must occur from further irradiation of the intermediate, $\text{cis-}[\text{Ru}(\text{biq})(\text{phen})(\text{CH}_3\text{CN})(\text{S})]^{2+}$ where $\text{S} =$ coordinating solvent molecule. It is also important to note that the growth of the ^1H NMR peak corresponding to free CH_3CN at 1.96 ppm for the irradiation of **2a** in CD_3CN shown in Fig. 4c does not begin until $t \sim 4$ min, which is consistent with the dissociation of CH_3CN taking place from the intermediate $\text{cis-}[\text{Ru}(\text{biq})(\text{phen})(\text{CH}_3\text{CN})(\text{CD}_3\text{CN})]^{2+}$ and not directly from **2a**.

In contrast to the results for **2a**, the $^{31}\text{P}\{\text{H}\}$ NMR of **1a** did not change as a function of irradiation time (Fig. S14†) in CD_3CN , indicating that the PPh_3 ligand is photostable in this complex and does not photodissociate. Together, these results demonstrate the photoinduced ligand exchange of PPh_3 from **2a** upon irradiation with $\lambda_{\text{irr}} \geq 395$ nm, while **1a** undergoes only CH_3CN ligand substitution. Steric strain around the $\text{Ru}(\text{II})$ center due to the bulky biq ligand is known to influence the exchange of ligands from $\text{Ru}(\text{II})$ complexes following irradiation and likely plays a role in the photoinduced PPh_3 exchange in **2a**.^{41,65,66} The observed trends highlight a need to further investigate the geometry around the ruthenium center in phosphine complexes to identify the origin of the unusual dissociation of PPh_3 in **2a**.

Structural comparisons

The generally accepted model for photoinduced ligand exchange in $\text{Ru}(\text{II})$ polypyridyl complexes is thermal population of a ^3LF (ligand field) state from a lower energy $^3\text{MLCT}$ (metal-to-ligand charge transfer) state, which places electron density on Ru-L orbital(s) with σ^* anti-bonding character, leading to ligand dissociation.^{67–70} Distortion in the pseudo-octahedral geometry around ruthenium metal lowers the energy of the e_g -type σ^* set, and consequently of the dissociative ^3LF state, leading to a decrease in the activation energy required to thermally populate it from the lowest energy $^3\text{MLCT}$ state and



increasing the efficiency of ligand exchange. The introduction of steric bulk *via* ligands containing methyl, phenyl, or quinoline moieties has been shown to sufficiently distort the octahedral geometry, resulting in an increase the quantum yield of ligand exchange.^{71–74}

In order to better understand if steric effects account for the differences in the photoreactivity of **1a** and **2a**, their solid-state single-crystal X-ray structures were determined and are shown in Fig. 8, along with relevant bond lengths and angles listed in Table 2 with additional X-ray data available in Tables S3 and S4.† The Ru–N bond lengths from the ruthenium center to the bidentate ligand *trans* to the PPh₃ ligand are 0.05 Å longer on average in **2a** (biq) as compared to **1a** (phen), respectively, indicating greater steric strain in the former. The greater steric hindrance in **2a** relative to **1a** is further supported by the Ru1–P1 bond length, which is 0.024 Å longer in **2a**, consistent with a weaker Ru–P bonding interaction in the biquinoline complex. The crystal structure of **2b** (Fig. 8), obtained by irradiating a solution of **2a** in CH₃CN with visible light overnight, possesses Ru–N3 and Ru–N4 bond lengths that are 0.075 and 0.028 Å shorter than in **2a**, respectively, where N3 and N4 are the nitrogen atoms in the biquinoline ligand. Such a decrease in Ru–N bond lengths demonstrates that the photodissociation of PPh₃ relieves steric strain.

Evidence for distortion of the bidentate ligand *trans* to PPh₃ is apparent in the torsion angle, the angle between the two N–C–C planes formed by N3 of the bidentate ligand and the two carbon atoms bridging it to N4, expected to be 0° in an ideal octahedral geometry. In the case of **1a**, the N3–C–C–N4 torsion angle in the phenanthroline ligand is 2(1)°. In contrast, a torsion angle of 10.5(2)° is measured in the biquinoline ligand in **2a**. Further, geometric planes defined by N1–Ru1–N2 and N3–Ru1–N4 would be at 90° angles in an ideal octahedral geometry and deviations from this angle reveal additional steric distortion around the metal center.⁴ The angle between these two planes in **1a** was determined to be 88.69°, which is reduced to 81.37° in **2a**. Importantly, in **2b** this angle is 86.93° and the N3–C–C–N4 torsion angle in the biquinoline ligand is 3.1(3)°, showing the substitution of PPh₃ for CH₃CN allows the complex

Table 2 Selected crystallographic bond lengths and angles for **1a**, **2a**, and **2b**

	1a	2a	2b
Bond lengths (Å)			
Ru1–N3	2.106(9)	2.148(1)	2.073(2)
Ru1–N4	2.06(1)	2.112(2)	2.084(2)
Ru1–N5	2.030(2)	2.036(2)	2.046(2)
Ru1–P1/N6	2.343(1)	2.3669(5)	2.034(2)
Torsion angles (°)			
N1–C–C–N2	1.1(4)	2.1(2)	0.6(3)
N3–C–C–N4	2(1)	10.5(2)	3.1(3)
Bond angles (°)			
N1–Ru1–N3	87.7(2)	94.09(9)	100.22(7)
N1–Ru1–P1/N6	93.68(8)	85.26(7)	82.24(7)
N2–Ru1–N3	91.2(2)	81.95(9)	87.52(6)
N3–Ru1–N5	82.1(2)	93.13(9)	93.52(7)
N4–Ru1–P1/N6	98.8(4)	103.63(7)	99.31(7)

to adopt a geometry closer to the ideal octahedral. It should be noted that the Ru–N5 bond to the CH₃CN ligand does not significantly differ in length (0.006 Å) in **1a** and **2a**.

The bond angles provided in Table 2 demonstrate additional differences in the steric distortion in **1a**, **2a**, and **2b**. For example, the N1–Ru1–N3 and N2–Ru1–N3 angles show the extent of distortion in the phenanthroline ligand in each complex. Substitution of biq for phen in **2a** pushes N1 towards the phosphine ligand while N2 moves away from P1 to accommodate the large PPh₃ unit, such that the N2–Ru1–N3 angle is nearly 10° greater in **1a** than **2a**. These same angles in **2b** demonstrate that the dissociation of PPh₃ relieves the steric strain between the polypyridyl bidentate ligands by allowing the phenanthroline ligand to move away from biq.

The angles N1–Ru1–P1 and N4–Ru1–P1 show the bulkier biquinoline ligand pushes the phosphine away from biq and toward phen in **2a**, to a significantly greater extent than in the analogous phen in **1a**. Similar to the bond angles, the CH₃CN ligand does not display great distortion. The only notable angle

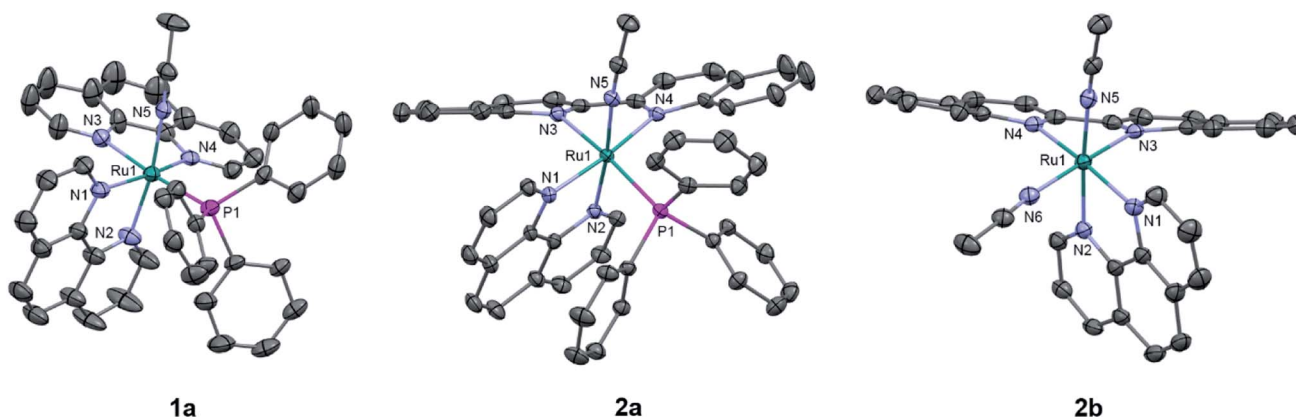


Fig. 8 ORTEP plots of **1a**, **2a**, and **2b** (thermal ellipsoids have been drawn at 50% probability and hydrogen atoms, PF₆[−] molecules, and co-crystallized solvent molecules have been omitted for clarity); Ru: cyan, N: light purple, C: grey, and P: magenta.



change is that for N3–Ru1–N5, which is 11° smaller in **1a** than in **2a**, indicating that PPh₃ pushes the acetonitrile ligand towards the *trans* bidentate ligand in **1a** but the presence of the larger biq ligand prevents this displacement in **2a**. In summary, the X-ray crystal structures show that there are significantly greater deviations from octahedral geometry in **2a** as compared to **1a** and **2b**.

Calculations

Density functional theory (DFT) calculations were performed to determine if the bonding and electronic structure in the complexes could further explain the differences in the photo-reactivity of **1a** and **2a**. Geometry optimizations in the singlet ground state (¹GS) of **1a**, **2a**, and **2b** resulted in structures in good agreement with experimental crystallographic data (Table S4). Table S4† shows that the calculated bond lengths, angles, and torsional angles are in good agreement to those obtained experimentally from the crystal structures of each complex. The ¹GS highest occupied molecular orbitals (HOMOs) in **1a** and **2a** exhibit primarily Ru-d orbital character, as is typical for Ru(II) polypyridyl complexes.^{63,75–77} The lowest unoccupied molecular orbital (LUMO) in each complex is primarily localized on the ligand *trans* to PPh₃, 1,10-phenanthroline in **1a** and 2,2'-biquinoline in **2a**; the latter agrees with the findings from electrochemistry (Fig. S15, S16 and Table S5†).

Geometry optimizations and vibrational frequency calculations were also performed in the triplet excited states (³ES) of **1a** and **2a**. In the ³ES, longer Ru–NCCH₃ and Ru–P bond distances are calculated in both complexes as compared to the corresponding ¹GS (Table 3). The calculated Ru–P bonds in the ³ES are similar, 2.475 Å in **1a** and 2.473 Å in **2a**, increasing from 2.411 Å and 2.441 Å in the ¹GS, respectively. However, the Ru–NCCH₃ bond is 0.04 Å longer in the ³ES of **1a**, 2.055 Å, as compared to that in **2a**, 2.015 Å, which may indicate that the Ru–nitrile bond is weaker in the excited state of **1a** relative to that in **2a**.

The differences in the bonds of **1a** and **2a** in the ³ES were further investigated by calculating the Mayer bond orders (MBOs) of the bonds involving ruthenium (Table 3). MBOs are an extension of Wiberg bond orders and can provide insight into the relative strengths of bonds in transition metal complexes. In the ³ES of **1a**, the MBO of the Ru–NCCH₃ bond

(Ru1–N5) exhibits a 26.6% decrease as compared to the ¹GS and the Ru–N(phen) bond *trans* to CH₃CN (Ru1–N3) displays a dramatic 74.0% increase. These results indicate that the reduced phenanthroline ligand in the excited state exerts a *trans*-type influence on the CH₃CN ligand and weakens the Ru1–N3 bond, likely contributing to its dissociation in the excited state as has been observed in other Ru(II) complexes.^{33,78,79} In contrast, the MBO of the bond to CH₃CN only decreases by 11.8% in the triplet state of **2a** and the order of the Ru–N bond *trans* to CH₃CN does not change in the triplet state of **2a**. In addition, the Ru–NCCH₃ bond itself is significantly stronger in the ³ES of **2a** as compared to **1a**, 0.576 and 0.441, respectively.

Both **1a** and **2a** displayed a ~6% increase in the MBO of the Ru–PPh₃ bond, Ru1–P1 in Table 3, in the ³ES, although the bond is weaker in the excited state of **2a**, MBO = 0.659, than in **1a**, MBO = 0.719. The order of the bond *trans* to the phosphine ligand, Ru1–N1, is calculated to increase by 31.0% in **2a** and 56.9% in **1a**, indicating a *trans*-type influence in both complexes. However, the phenomenon is significantly stronger in **1a** for the bond positioned *trans* to the CH₃CN ligand, Ru1–N3, which may explain why the CH₃CN ligand preferentially photodissociates in this complex.

Mulliken spin density (MSD) calculations were also performed on the lowest energy ³ES of **1a** and **2a**. These calculations determine the unpaired electron density on each atom in the ³ES and can provide information on the nature of the excited state. In complexes with ³LF as the lowest energy triplet excited state, the spin density on the Ru(II) metal center would theoretically equal two. If the lowest energy ³ES is MLCT in nature, the spin density on ruthenium is expected to be one, and any deviation from these whole numbers indicates metal/ligand mixing. The MSD on ruthenium and the summed density on each ligand in the lowest energy ³ES of **1a** and **2a** is shown in Fig. 9. The calculated spin densities on Ru(II) indicate the lowest energy triplet excited state is MLCT in nature with notable ligand character in **1a** and significant mixing from a ligand-centered state in **2a**. The summed spin density on the phenanthroline ligand *trans* to PPh₃ in **1a** is 0.447, lower than the 0.763 sum on the phen ligand *trans* to CH₃CN, further indicating stronger *trans*-type influence on the nitrile ligand in the excited state. In the ³MLCT of **2a** the sum of the spin density on the phen ligand, which is *trans* to CH₃CN, is 0.074 and a sum of 1.433 is calculated on the biquinoline ligand, indicating a significant *trans*-type influence on the PPh₃ ligand. Taken together with the calculated bond lengths and MBOs in the ³MLCT state, it can be concluded that while **1a** exhibits a *trans*-type influence on both monodentate ligands, it is stronger in the case of the CH₃CN ligand which can explain the preferential dissociation of CH₃CN upon irradiation. In the ³MLCT state of **2a** there is no evidence of *trans*-type influence on the nitrile ligand, but a significant degree of *trans*-type influence is calculated for the phosphine ligand. These results, in conjunction with the steric strain evident from the crystal structure, can explain the unexpected photodissociation of the PPh₃ ligand in **2a**.

Table 3 Selected Mayer bond orders, MBOs, and calculated bond lengths in the ¹GS and lowest ³ES of **1a** and **2a**

Bond	MBOs				Bond Lengths/Å			
	1a		2a		1a		2a	
	¹ GS	³ ES	¹ GS	³ ES	¹ GS	³ ES	¹ GS	³ ES
Ru1–N3 ^a	0.370	0.656	0.223	0.223	2.080	2.022	2.096	2.095
Ru1–N5 ^b	0.601	0.441	0.653	0.576	2.008	2.055	2.000	2.015
Ru1–N1 ^c	0.260	0.408	0.281	0.368	2.136	2.103	2.170	2.146
Ru1–P1	0.760	0.719	0.701	0.659	2.411	2.475	2.441	2.473

^a Bond to N atom of bidentate ligand *trans* to CH₃CN. ^b Bond to N atom of CH₃CN. ^c Bond to N atom of bidentate ligand *trans* to PPh₃.



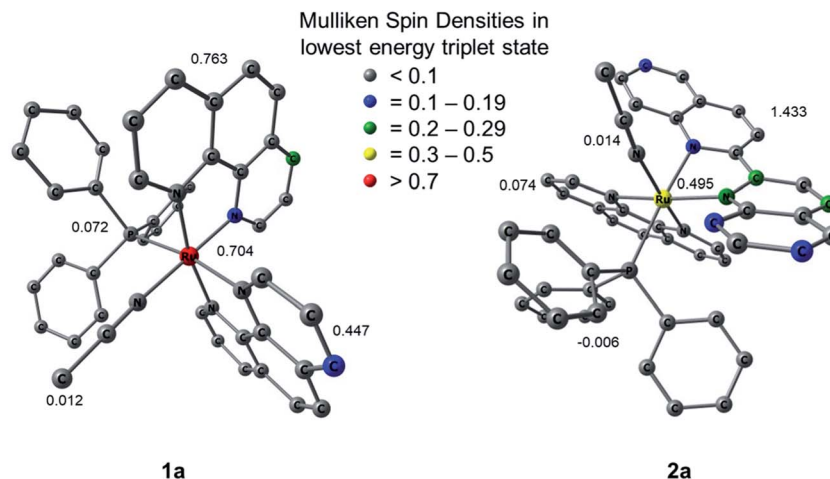


Fig. 9 Mulliken spin densities (MSDs) on ruthenium and the summed densities on each ligand in the calculated lowest energy triplet excited states of **1a** and **2a**.

Cell viability

The effect of photochemical reactivity on the biological behavior of **1a** and **2a** was investigated on the MDA-MB-231 triple-negative breast cancer cell line in the dark and upon irradiation and the results were compared to those of related complexes previously reported.^{18,80,81} The half effective concentration, EC_{50} , was determined, which is defined as the concentration of the compound when the viability of the experimental cells is 50% compared to those in the absence of active compound. Complexes **1a** and **2a** were incubated with MDA-MB-231 cells for 1 h and were then either left in the dark or irradiated with blue light for 20 min ($\lambda_{irr} = 460\text{--}470\text{ nm}$, 56 J cm^{-2}). Cellular viability was assessed using the MTT assay after 72 h, where the viability upon treatment with only the vehicle (1% DMSO) only in DMEM (Dulbecco's Modified Eagle Medium) was considered as 100%, and the results are listed in Table 3. Both **1a** and **2a** are non-toxic against MDA-MB-231 cells in the dark with EC_{50} values of $26.6 \pm 1.5\ \mu\text{M}$ and $>30\ \mu\text{M}$, respectively; the latter is above the maximum concentration allowed by the solubility in growth media ($30\ \mu\text{M}$). However, irradiation with blue light significantly increased the toxicities of both complexes, resulting in $EC_{50}^L = 4.6 \pm 0.6\ \mu\text{M}$ for **1a** and $EC_{50}^L = 7.1 \pm 0.2\ \mu\text{M}$ for **2a** (Fig. S17 and S18†). These values of the phototherapeutic index, PI, values defined as EC_{50}^D/EC_{50}^L , were calculated to be 5.8 and >4.2 for **1a** and **2a**, respectively (Table 4). While complex **1a** exchanges CH_3CN for a solvent water molecule, **2a** releases PPh_3 upon irradiation. The finding that PPh_3 by itself is not toxic against MDA-MB-231 cells both in the dark and when irradiated with blue light ($t_{irr} = 20\text{ min}$, $\lambda_{irr} = 460\text{--}470\text{ nm}$, 56 J cm^{-2}) in the concentrations used for complexes **1a** and **2a** (Fig. S19†) led to the conclusion that the corresponding aqua complexes are the major cause of toxicity upon irradiation.

Comparison of the cell toxicity data of $cis\text{-[Ru(bpy)}_2(\text{PPh}_3)(\text{CH}_3\text{CN})]^{2+}$ and $cis\text{-[Ru(bpy)}_2(\text{CH}_3\text{CN})_2]^{2+}$ in Table 3 indicates that the substitution of a CH_3CN ligand for PPh_3 increases the PI value of the ruthenium complex. Complexes **1a** and **2a** exhibit PI values similar to or exceeding that of $cis\text{-[Ru(bpy)}_2(\text{PPh}_3)(\text{CH}_3\text{CN})]^{2+}$. These results indicate that the phosphine ligand

Table 4 EC_{50} values upon irradiation and in the dark and phototherapeutic index, PI, for **1a**, **2a**, and related compounds in MDA-MB-231 cancer cells

Complex	$EC_{50}^D/\mu\text{M}^a$	$EC_{50}^L/\mu\text{M}^a$	PI ^b
1a	26.6 ± 1.5	4.6 ± 0.6	5.8
2a	>30	7.1 ± 0.2	>4.2
$[\text{Ru}(\text{bpy})_2(\text{PPh}_3)(\text{CH}_3\text{CN})]^{2+ c}$	>30	7.0 ± 1.4	>4.3
$cis\text{-[Ru}(\text{bpy})_2(\text{CH}_3\text{CN})_2]^{2+ d}$	244 ± 23	223 ± 94	1.1 ± 0.4

^a Data are an average of three independent experiments. ^b $PI = EC_{50}^D/EC_{50}^L$. ^c From ref. 82; bpy = 2,2'-bipyridine. ^d From ref. 83; IC_{50} values against HeLa cells.

may positively influence the activity against cancer cells of these ruthenium complexes, a conclusion supported by previous examples of phosphines improving cellular uptake and localizing complexes to the mitochondria, increasing selectivity for cancerous cells over healthy cells,²² and increasing cytotoxic activity against breast and colon cancer cells of phosphine containing *cis*-configured Pt(II) complexes over that of cisplatin.^{21,84–86}

Conclusion

Two new triphenylphosphine-containing complexes, **1a** and **2a**, were synthesized and their ground state spectroscopic and electrochemical properties were characterized, along with their photoinduced ligand exchange and cytotoxicity against a triple-negative breast cancer cell line. Changes in the electronic absorption and NMR spectra of complex **1a** revealed the substitution of a CH_3CN ligand for a solvent molecule following visible light irradiation and a substitutionally inert PPh_3 . In contrast, the photolysis of complex **2a** results in the initial exchange of the PPh_3 ligand generating a solvated intermediate, and the latter goes on to absorb a second photon which then undergoes CH_3CN substitution. A comparison of the single crystal X-ray structures reveals that **2a** exhibits greater steric



distortion around the metal center than **1a**, which is subsequently relieved upon the photoinduced exchange of PPh₃ for a less sterically-demanding solvent molecule. To our knowledge, this represents the first report of the photodissociation of a phosphine ligand from a Ru(II) polypyridyl complex. In addition, the ability of phosphine ligands to enhance cellular uptake was shown to enhance the photocytotoxicity of **1a** and **2a** against a triple-negative breast cancer cell line relative to related complexes without PPh₃ in their coordination sphere. This work shows that a coordinated PPh₃ ligand can serve as a new architecture for potential therapeutics for use in PCT.

Author contributions

Sean J. Steinke synthesized the complexes and conducted photophysical and photochemical measurements. Eric J. Piechota aided with data analysis and interpretation, Sayak Gupta performed the cell studies, and Curtis E. Moore collected and solved the crystal structures. Jeremy J. Kodanko and Claudia Turro advised their respective group members.

Conflicts of interest

The authors declare no competing financial interest.

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