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Two-dimensional materials for electrochemical CO₂ reduction: materials, *in situ/operando* characterizations, and perspective

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Electrochemical CO₂ reduction (CO₂ ECR) is an efficient approach to achieving eco-friendly energy generation and environmental sustainability. This approach is capable of lowering the CO₂ greenhouse gas concentration in the atmosphere while producing various valuable fuels and products. For catalytic CO₂ ECR, two-dimensional (2D) materials stand as promising catalyst candidates due to their superior electrical conductivity, abundant dangling bonds, and tremendous amounts of surface active sites. On the other hand, the investigations on fundamental reaction mechanisms in CO₂ ECR are highly demanded but usually require advanced *in situ* and *operando* multimodal characterizations. This review summarizes recent advances in the development, engineering, and structure–activity relationships of 2D materials for CO₂ ECR. Furthermore, we overview state-of-the-art *in situ* and *operando* characterization techniques, which are used to investigate the catalytic reaction mechanisms with the spatial resolution from the micron-scale to the atomic scale, and with the temporal resolution from femtoseconds to seconds. Finally, we conclude this review by outlining challenges and opportunities for future development in this field.

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1 Introduction

The historically unprecedented atmospheric CO₂ levels of current times have resulted in global warming and extinction risks to many animal and plant species.^{1,2} Therefore, the fixation and conversion of CO₂ is critical in realizing energy and

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environmental sustainability. Over the past decades, many CO₂ reduction reaction (CO₂RR) strategies have been developed to produce high-value fuels or commodity chemicals to address this problem. Among these approaches, the electrochemical CO₂ reduction (CO₂ ECR) is a well-studied strategy that employs electrochemical electrolyzers to convert CO₂ to value-added fuels using renewable energy and electrochemical catalysts. Compared to other CO₂ conversion approaches, CO₂ ECR has attracted great attention due to many of its advantages: (1) mild work environment at room temperature and atmospheric pressure, (2) the ability to yield various high-value-added chemicals, especially HCOOH, CO, CH₄, CH₃OH, CH₃CH₂OH, C₂H₄, *etc.* (3) tunable chemical reactions – parameters such as electrolytes, operating voltage, and reactants can be easily adjusted based on the desired reactions, (4) the potential to utilize renewable energy as the energy source and industrial waste as the reactants, and (5) variety of available options for reactor designs.^{3–5} However, in CO₂RR, molecule CO₂ is chemically inert that the activation of the C=O bond in CO₂ requires high overpotentials and high energy at 750 kJ mol⁻¹.⁶ Further complicating the reaction, the highly intricate electrochemical CO₂ reduction processes occur in an aqueous environment and involve multiple electrons and protons transfer in various reaction paths, producing a wide variety of chemical products.^{6,7} To this day, the mass production of multicarbon CO₂RR products, such as ethanol and ethylene, is still challenging and far from real-life applications, as only the catalysts with superior morphological and electronic structure are capable of achieving such highly demanded reactions.^{6,8,9} The adsorption, interaction, and desorption of CO₂ and multiple reaction intermediates on catalyst surfaces are the key steps in CO₂ ECR processes and are determined by such catalyst properties.¹⁰ Therefore, engineering low-cost, green electrocatalysts with excellent durability, high activity, and desired selectivity is crucial in developing the next-generation CO₂ reduction reaction (CO₂RR) applications.

The demand for high-efficiency catalysts for electrochemical CO₂RR has stimulated tremendous research efforts in the field.⁹ Over the past decades, the research focus shifted

from bulk materials to nanomaterials and then from nanomaterials to advanced two-dimensional materials. Early works on the subject focused on metal bulk materials such as Cu, Au, Pd, and Ag, which all demonstrated the catalytic activities of such catalysts to convert CO₂ into hydrocarbons, CO, methanol, *etc.* However, the activity, durability, and selectivity of these catalysts are low for CO₂ reduction to meet practical demand.^{11–15} To address these shortcomings, the rapid development of CO₂RR nanocatalysts has improved the performance of CO₂RR in the catalytic activity, selectivity, and stability due to a large amount of surface active sites of nanocatalysts.^{16–20} However, the current state-of-the-art systems are still far from practical applications.

Ultra-thin two-dimensional (2D) catalysts with atomic-scale thicknesses have demonstrated outstanding capabilities and performance as CO₂ ECR catalysts. Compared with nanoparticles and their bulk materials, 2D materials are excellent catalyst candidates, owing to their intrinsic advantages such as ultrahigh specific surface area, abundant exposed surface atoms, atomically flat facet, and high electrical conductivity.^{3,21,22} Many structural, chemical, and electronic properties can be fine-tuned based on the desired reaction paths and products *via* various strategies, including defect engineering, doping, surface structural modification, and surface functionalization.^{23–27} On the other hand, the study of 2D catalysts for CO₂RR can also help us understand the relationships between active sites and the catalysis mechanism and performance at the atomic scale.³ Hence 2D solids are usually considered as an ideal model catalyst to study CO₂RR mechanisms. The emergence of 2D materials offers a promising opportunity for the further development and applications of electrochemical CO₂ reduction.

Meanwhile, *in situ* and *operando* characterizations have become the essential tools to understand CO₂ RR mechanisms and evaluate catalyst performance. The electrocatalytic reductions of CO₂ are extremely complicated: multiple reaction pathways can lead to the generations of various reaction intermediates and products, requiring the transfer of multiple protons and electrons. During CO₂ reduction reactions,



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advanced *in situ* and *operando* characterizations can monitor the generations of both reaction intermediates and products, identify catalytic active sites under the real time, probe the state and properties of catalysts, and reveal the reaction pathways in dynamic reaction environment.^{28,29} While many such advanced techniques have been developed and widely utilized for CO₂ ECR studies, new concepts and techniques have attracted much attention for the discoveries of novel catalysts and studies of fundamental reaction mechanisms.

In this review, we summarize the research progress on four types of representative 2D materials employed in electrochemical CO₂RRs: (1) transition metals (TMs), (2) transition metal oxides (TMOs), (3) transition metal dichalcogenides (TMDs), and (4) other types of 2D materials (MXenes, metal organic framework, covalent organic frameworks). Then, we describe current regulating strategies and structure–activity correlations of 2D materials for electrocatalytic CO₂RR. In addition, several state-of-the-art *in situ* and *operando* characterization techniques are reviewed. Finally, this review will conclude with a perspective on upcoming challenges, opportunities, and trends of 2D materials of future 2D electrochemical CO₂ reduction catalysts research.

2 Nanostructured two-dimensional materials for CO₂ electroreduction

To date, numerous two-dimensional (2D) materials have been developed and utilized in CO₂ ECR applications, leading to many exciting and promising findings and material candidates. In general, there are five categories of 2D materials that have demonstrated promising CO₂ ECR performance: a. 2D transition metals, b. 2D transition metal oxides, c. 2D transition metal dichalcogenides (TMDs), d. 2D carbon-based materials, and e. metal–organic frameworks and covalent

organic frameworks. Fig. 1 shows representative transmission electron microscopy (TEM) and atomic force microscopy (AFM) images of these five categories of 2D materials reviewed in this article.

2.1 Two-dimensional 2D transition metals (TMs)

Transition metals have been extensively studied for CO₂RR for decades.³⁵ Notably, transition metal nanosheets are highly effective electrocatalysts for CO₂ ECR due to their abundant active sites and high conductivity.³⁶ Such advantages can benefit CO₂ ECR *via* different mechanisms. By optimizing the electronic properties, Bi nanosheets have exhibited high selectivity towards formic acid in CO₂RR and a high faradaic efficiency of 92% in 0.5 M KHCO₃ solution at ~1.5 V (*versus* SCE). Such 2D nanosheets also generated lower onset potential and about five times higher current density compared to commercial Bi nanoparticles.³⁷ Density functional theory (DFT) calculations proved that 10 nm-thick Bi nanosheets have a higher density of states (DOS) at the Fermi level than bulk Bi, leading to improved CO₂RR performance with high selectivity and excellent stability for electrocatalysis CO₂ reduction.³⁸ The morphological advantages of 2D TMs have also enhanced CO₂ ECR performance. DFT calculations for Bi nanosheets synthesized *via* a molten-salt-assisted method suggested that the edge sites of Bi nanosheets demonstrate superior activity than terrace sites.³⁹ In terms of surface properties, Bi nanosheets obtained by liquid-phase exfoliate method exhibited greater hydrophilicity than Bi bulk materials, thus promoting the electron transfer between solid and liquid. These Bi nanosheets produced a high current density of 16.5 mA cm⁻² at -1.1 V (*versus* RHE), a faradaic efficiency of 86.0% for formate production, and well-preserved catalytic activity after 10 h of testing.³⁶ In addition, advanced 2D structures have shown promising performance in CO₂ ECR. 2D mesoporous Bi nanosheets prepared *via* the cathodic conversion method



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Fig. 1 An overview of different types of 2D materials. (a) and (b) TEM and AFM images of 2D palladium–copper alloy nanodendrites; (c) SnO₂ modified two-dimensional CuO nanosheet; (d) TEM image of MoS₂ nanosheet; (e) TEM image of Ultrathin 2D Mxene. (f) and (g) TEM images of Bi-MOF ultrathin nanosheet.^{30–34}

demonstrated a large current density of $\sim 18 \text{ mA cm}^{-2}$ at -1.1 V , a high faradaic efficiency for formate conversion of almost 100%, and an excellent durability of 12 h.⁴⁰ A recent report of hydrangea-like superstructure consisting of ultrathin Bi nanosheets demonstrated a faradaic selectivity of over 95% with a wide potential window ranging from -0.78 V to -1.18 V and a high current density of over 300 mA cm^{-2} .⁴¹

In a recent report, 50 nm-thick Ag nanosheets with hierarchical structures demonstrated a large surface area for catalytic reactions and gas transport/diffusion, resulting in a high faradaic efficiency of 95% for CO₂ conversion into CO and a

current density that is 37 times higher than that of Ag nanoparticles at -0.6 V (*versus* RHE).⁴² Multilayered Zn nanosheets also showed high faradaic efficiency of $\sim 86\%$ for CO product at 1.13 V (*versus* RHE) and good stability during a 7-hour test. The high density in edge sites in Zn nanosheets resulted in a density 9 times higher than that of bulk Zn foil.⁴³ Ultrathin copper nanosheets prepared by galvanic replacement method also exhibited high selectivity and activity for CO₂ reduction, where the optimized hierarchical Cu electrodes possessed a faradaic efficiency of 74.1% for CO product and a current density of 23.0 mA cm^{-2} at -1.0 V (*versus* RHE).⁴⁴ Abundant vacancy defects in ultrathin Cu nanosheets can improve the initial kinetics of CO formation and accelerate the mass and electrons transport in the CO₂ reduction process. Moreover, hybrid Cu/Ni(OH)₂ nanosheets, which consist of atomically thick Cu nanosheets loaded on Ni(OH)₂, also displayed excellent activity, selectivity, and durability (22 hours) in electrocatalytic CO₂ reduction.⁴⁵ The nanosheets provided a large current density of 4.3 mA cm^{-2} at -0.5 V (*versus* RHE) and a high faradaic efficiency of 92% for CO product.⁴⁵ Ultrathin 2D hierarchical Pd/SnO₂ nanosheets, namely Pd nanosheets partially capped by SnO₂ nanoparticles, displayed a maximum faradaic efficiency (FE) of $\sim 54.8\%$ at -0.24 V (*versus* RHE) and high stability for electroreduction of CO₂ into CH₃OH.⁴⁶ The enhanced performance of Pd/SnO₂ nanosheets could be attributed to the abundance of Pd–O–Sn interfaces. Notably, recent advances in 2D alloys have demonstrated enhanced CO₂ ECR performance than their single-component counterparts. A palladium–copper alloy nanodendrite system exhibited a high Formate selectivity of over 90% at a low cathodic over potential of $-0.4 \text{ V}_{\text{RHE}}$.³⁰ Fig. 1a and b demonstrate the morphology of this nanodendrite system. DFT calculations revealed that Cu



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atoms can weaken the *CO adsorption to avoid CO surface poisoning and promote the *OCHO adsorption on neighboring Pd atoms.³⁰ Such synergistic effects shall inspire more research endeavors in 2D alloy systems in the future.

Ultrathin metal nanosheets are prone to oxidation, which leads to a rapid decay in cyclability. However, many studies have also demonstrated through theory calculations and various characterizations that MO_x layer on the surface of transition metals plays a vital role in electrocatalytic CO₂.^{47–49} Recently, Xie *et al.* successfully prepared 4-atoms-thick Co nanosheets (~0.84 nm) with different surface structures by a ligand-confined growth strategy.⁵⁰ Partially oxidized Co nanosheets showed an onset potential of –0.68 V for formate production in 0.1 M Na₂SO₄ solution, a high current density of 10.59 mA cm^{–2}, a faradaic efficiency of 90.1% towards formate, and durability of over 40 h.⁵⁰ Furthermore, catalytic activity for these 4-atom-thick Co nanosheets showed a 260-fold improvement compared to bulk metal Co, stemming from a large number of surface active sites in ultrathin 2D nanosheets and the existence of surface Co oxides. Partially oxidized Co nanosheets with a smaller Tafel slope of ~59 mV dec^{–1} show a fast pre-equilibrium for the formation of intermediate CO₂^{•–}, suggesting that surface Co oxide in nanosheets can accelerate CO₂ activation and facilitate the rate-determining step for H⁺ transfer in chemical reaction.⁵⁰

2.2 Two-dimensional 2D transition metal oxides (TMOs)

Transition metal oxides have been extensively studied for usage in photoelectrochemical and electrochemical CO₂ reduction reactions because of their controlled structure, tunable energy band gap, and low cost. Their unfilled orbitals and active d electrons can energetically facilitate the bonding between metal atoms and CO₂ molecules.⁵¹ However, traditional bulk transition metal oxides' low specific surface area and poor electrical conductivity are two key drawbacks that limit their activity and stability in CO₂ reduction.⁵² Therefore, the emergence of two-dimensional (2D) transition metal oxides provides a promising opportunity to develop CO₂ reduction reaction. It has been reported that 3D hierarchical SnO₂ nanosheets, at a moderate overpotential of 0.88 V, demonstrated a high faradaic efficiency of 87% for the production of formate and a partial current density of 45 mA cm^{–2}.⁵³ This outstanding performance stemmed from the large surface area and fast electron transfer in 2D nanosheets. 1.72 nm-thick Co₃O₄ nanosheets prepared by solvothermal reaction also exhibited high activity for electrocatalytic CO₂ reduction that the current density reached 0.68 mA cm^{–2} at –0.88 V (*versus* SCE).⁵³ The formate faradaic efficiency was over 60% in 20 h.⁵² Based on this, engineered V-rich single-unit-cell layer Co₃O₄ possessed higher current densities of 2.7 mA cm^{–2} and 85% selectivity for formate product at –0.87 V (*versus* SCE).⁵⁴ DFT calculations and experiments demonstrated that abundant oxygen vacancies could promote the adsorption of CO₂ on Co₃O₄ and the desorption of HCOO[–], while further lowering the rate-limiting activation energy, leading to improvement in activity and durability in the cata-

lytic reaction. Some other researchers also found that ZnO nanosheets with rich oxygen vacancies display a high current density of –16.1 mA cm^{–2} at –1.1 V (*versus* RHE) and a high faradaic efficiency of 83% for CO production compared with ZnO nanosheets and ZnO nanosheets with poor oxygen vacancies.⁵⁵ DFT calculations and various characterizations clarified that large surface specific area and rich oxygen vacancies could increase surface active sites and facilitate the activation of CO₂ molecular.⁵⁵ Up to now, many other ultrathin transition metal oxide nanosheets made of TiO₂,⁵⁶ WO₃,⁵⁷ Bi₂WO₆,⁵⁸ ZnGa₂O₄,⁵⁹ SnNb₂O₆,⁶⁰ or BiVO₄⁶¹ have been successfully prepared *via* a variety of methods and these 2D materials all also exhibited excellent performance for photocatalytic CO₂ reduction. Some optimization strategies, such as doping, have also demonstrated enhancing effect on CO₂ ECR performance using 2D metal oxide catalysts. By modifying 2D CuO with SnO₂, Lan *et al.* recently reported a high faradaic efficiency of 22% for C₂H₄ at a low over potential of –1.0 V_{RHE}.³¹ The outstanding performance was attributed to the facilitated CO₂-substrate bonding on Sn modified CuO. Fig. 1c shows the morphology of the catalyst studied in the research.³¹

2.3 Two-dimensional 2D transition metal dichalcogenides (TMDs)

Transition-metal dichalcogenides (TMDs) are a promising class of two-dimensional materials which are broadly applied in electrocatalytic reactions because of their unique metal-sulfur covalent bond, atomically flat facet, and high specific surface area.²¹ Notably, WS₂ and MoS₂ are two representative 2D TMDs electrocatalysts for CO₂ reduction, and the main products in such systems are CO chemicals.⁶² A representative TEM image of MoS₂ nanosheet is shown in Fig. 1d.³² To note, the use of 2D TMDs electrocatalysts, unlike other types of materials, requires an ionic liquid (1-ethyl-3-methylimidazolium tetrafluoroborate, EMIM-BF₄) in the electrolyte. The ionic liquid plays a crucial role in forming the [EMIM-CO₂]⁺ complex in acidic environments, which could effectively suppress the HER side reaction.^{62,63} Abbasi *et al.* demonstrated that metal-terminated edges in TMDs are mainly responsible for the remarkable dominated catalytic activity in CO₂ reduction due to high density of d-electrons and metallic character *via* DFT calculations. A high spin-polarized partial DOS for Mo atoms at edge sites of MoS₂ could improve the catalysis efficiency of CO₂ reduction, and the rate-determining step in CO₂ reduction is the desorption of CO* species adsorbed on edge Mo atom of MoS₂. It is found that the electron-transfer properties in 2D transition metal dichalcogenides determined the catalytic activity of CO₂ reduction.⁶⁴ Doping 5% Nb element into MoS₂ could facilitate the shift of the d-band center of edge Mo atom toward a less negative value, leading to a weakened binding of CO and 1 order improvement for CO turnover frequency. Asadi *et al.* evaluated the CO₂ electrocatalytic activities of a series of 2D TMDs, including MoS₂, WS₂, MoSe₂, and WSe₂ nanoflakes and revealed that the electron transfer properties of TMDs are critical in enhancing their CO₂ reduction performances.⁶³ With

the lowest work function among all nanoflakes prepared in this study, WSe₂ nanoflakes terminated with W atoms exhibited the excellent performance in the conversion of CO₂ into CO, showing a high current density of 18.95 mA cm⁻², a faradaic efficiency of 24%, and a turnover frequency of 0.28 s⁻¹ for CO formation at -0.164 V (*versus* RHE).⁶³ Furthermore, ternary TMD alloys demonstrated enhanced CO₂ reduction performance than their binary TMD counterparts. Notably, for CO production performed at -1.15 V (*versus* RHE), MoSeS alloy monolayers displayed a higher faradaic efficiency (45.2%) for CO production than MoS₂ monolayers (16.6%) and MoSe₂ monolayers (30.5%), and its current density (43 mA cm⁻²) was about 1.3 and 2.7 times higher than that of MoSe₂ and MoS₂ monolayers, respectively.⁶⁵ This improvement is attributed to the optimized d-band electronic structure of Mo atoms in alloys. Such optimization increases density of states near conduction band edge and improves electron transfer.⁶⁵ Theoretical calculations further illustrated that the off-center charge around Mo atoms not only better stabilized the COOH* intermediate, but also liberated the rate-limiting CO desorption step.⁶⁵ In this aspect, Nørskov and coworkers found that breaking the linear scaling relations between the binding energies of intermediates could enhance the catalytic efficiency of CO₂ reductions and promote the stabilization of intermediates, such as COOH*, CHO*, COH*, and CO*.⁶⁶ By theoretical calculations, they demonstrated that doped MoS₂ can create two binding sites, namely metal and sulfuring binding sites, and thus two different linear scaling relationships.⁶⁶ The binding of CO and CHO* on the sulfur site can form different moieties: a close-shell “SCO” moiety for CO and an open-shell “SCHO” moiety for CHO*. “SCO” moiety weakens the neighboring S–Ni bond and increases the distortion, while “SCHO” moiety showed little impact on the neighboring S–Ni interactions and minimal distortion energy. This discovery should also fit other metal/p-block materials, where two binding sites are present.

Notably, many other ultrathin two-dimensional transition metal dichalcogenides have been applied in photocatalytic CO₂ reductions, which offers an important opportunity for photoelectrocatalytic CO₂ reduction. One-unit-cell ZnIn₂S₄ layers with poor zinc vacancies and with rich zinc vacancies were prepared *via* hydrothermal method to study photocatalytic CO₂ reduction.⁶⁷ ZnIn₂S₄ with rich zinc vacancies exhibited 3.6 times higher than ZnIn₂S₄ with poor vacancies in formate yield, and its catalytic activity remain unchanged after 24 h photocatalysis.⁶⁷ Whereafter, atomically thin CuIn₅S₈ and that with sulfur deficient were synthesized to measure photocatalytic CO₂ reduction. 2D CuIn₅S₈ generally contained poor selectivity for CO and CH₄ from CO₂, while those with sulfur deficiencies showed higher activity and selectivity. Experimental results demonstrated that charge-enriched Cu–In dual sites in sulfur-deficient CuIn₅S₈ could promote the formation of stable Cu–C–O–In intermediate, which made it possess near 100% selectivity towards photocatalytic production of CH₄ from CO₂.⁶⁸ This regulation not only changed the reaction pathway to form CH₄ instead of CO in CO₂ reduction but also decreased the overall activation energy

barrier in catalytic reaction. Furthermore, 2D SnS₂ nanosheets were used to study photocatalytic CO₂ reduction. Partially oxidized SnS₂ nanosheets showed excellent catalytic activity for CO production, which near 2.3 and 2.6 times higher than the poorly oxidized SnS₂ nanosheets and the SnS₂ nanosheets.⁶⁹ These works confirmed the potential application of untapped two-dimensional transition metal dichalcogenides in photoelectrocatalytic CO₂ reduction, which may also achieve high-efficiency conversion of important C₂ production.

2.4 Two-dimensional 2D carbon-based materials

Two-dimensional carbon-based materials, such as graphene and g-C₃N₄, have attracted intense attention in CO₂ reduction due to their outstanding electrical conductivity, low cost, and adjustable surface structure.⁷⁰ Despite these qualities, two-dimensional pristine graphene is relatively chemically inert, and it has a weak absorption for molecular CO₂ and poor catalytic activity for electrochemical CO₂ reduction. The well delocalized π bonding network in pristine graphene possesses a high activation energy barrier for the endothermic absorption of key intermediates (*e.g.*, COOH*, OCHO*).⁷¹ Specifically, the undoped zigzag edge of graphene displays a relatively high energy barrier of \sim 1.3 eV for CO₂ absorption in electrochemical CO₂ reductions.⁷² Therefore, engineering the electronic structure of graphene is of vital importance to improve electrocatalytic activity. The enhancement approaches will be discussed in later sections.

MXenes, a category of emerging two-dimensional solid materials made up of transition metal carbides, carbonitrides, and nitrides, received much attention in the field of electrocatalysis.⁷³ The general formula of MXene is M_{n+1}X_nT_x ($n = 1, 2, 3$), where M is a transition metal, X represents carbon and/or nitrogen, and T denotes the surface terminations, such as O, OH, and F.⁷⁴ MXenes layers are attached together *via* van der Waals forces. MXenes materials are generally prepared by selective etching of specific atomic layers in layered precursors (MAX phases), and their surface contains various functional group after treatment. Therefore, the surface termination in MXenes has significant impacts on the electronic and catalytic properties.^{25,75} Li *et al.* first predicted the outstanding CO₂ capture capability of two-dimensional MXenes from group IV to VI series through density functional theory calculations. Furthermore, they found that the Cr₃C₂ and Mo₃C₂ MXenes possess high selectivity and activity for electrocatalytic CO₂ reduction, and these surfaces are more inclined to chemisorbing CO₂ than H₂O.⁷⁶ In the conversion process of CO₂ \rightarrow CH₄, the formation of OCHO* and HOCO* radical species is a part of the spontaneous reactions in the early hydrogenation steps, and it is considered the rate-determining step for CO₂ reduction. DFT results showed that MXenes (Cr₃C₂ and Mo₃C₂) with surface functional groups (*e.g.*, –O or –OH) can dramatically decrease the overpotentials in electrocatalytic CO₂ reduction compared to bare Cr₃C₂ and Mo₃C₂, reducing the energy input from 1.05 and 1.31 eV to 0.35 and 0.54 eV for Cr₃C₂ and Mo₃C₂, respectively.⁷⁶ Chen *et al.* systematically studied electrocatalytic activity for CO₂ reduction reaction on

17 different MXenes with –OH termination groups *via* the first-principles approach.²⁵ Among these catalysts, Sc₂C(OH)₂ and Y₂C(OH)₂ showed the least negative limiting potential of –0.53 and –0.61 V in catalyzing CO₂ reduction into CH₄.²⁵ Furthermore, they found that the adsorbed species can be better stabilized by the H atoms in –OH[–] terminated MXenes and that the strong –H binding can suppress the competing HER reaction. Such mechanisms can dramatically improve the activity, selectivity, and stability of electrocatalytic CO₂ reduction. Bader charge analyses confirmed that less charge migration in MXene catalysts during the potential-limiting step (*H)COOH → *CO elementary step) can well promote the catalytic performance for CO₂RR.²⁵

Handoko *et al.* also used DFT calculations to predict the property of two-dimensional MXenes in electrocatalytic CO₂ reduction, where W₂CO₂ and Ti₂CO₂ were found to be the most promising MXene catalysts with low overpotential and good selectivity for CO₂ reduction to CH₄.⁷⁷ DFT results showed that the accessible *HCOOH pathway is more favorable than the *CO route in electrocatalytic CO₂ reduction process and that the binding energies of *COOH and/or *HCOOH are the important limiting step of MXene catalysts.⁷⁷ Furthermore, their study also showed that O-termination functional groups on MXenes have vital effects on stabilizing the reaction intermediates and are the key reasons why MXenes possess good activity and high selectivity for CO₂ reduction. Soon afterwards, Xiao *et al.* also systematically investigated the CO₂ reduction reaction in 2D MXenes (M = Sc, V, Cr, Mn, Zr, Nb, Mo, Mo₂Ti, Hf, Ta, and W) *via* DFT calculations.⁷⁸ They found that the transition metal in catalysts has important effects on the active site of CO₂ reduction and that the surface bicarbonate species on MXenes can suppress HER reaction for improving the selectivity of CO₂RR. Furthermore, two possible catalytic mechanisms for CO₂ reduction into CH₄ were also proposed:⁷⁸

Mechanism I:



Mechanism II:



Among these MXenes, Sc₃C₂, Ta₃C₂, and Ti₃C₂ are considered efficient electrocatalysts for reducing CO₂ to CH₄. Other materials, such as Mo₃C₂, V₃C₂, Mo₂TiC₂, and W₃C₂, are predicted to possess lower overpotentials of 0.74 V, 0.45 V, 0.17 V, and 0.41 V in CO₂RR, respectively.⁷⁸ In addition, single atom catalyst systems have also been studied with MXene. Recently, single atom copper immobilized MXene Ti₃C₂Cl_x demonstrated a high faradaic efficiency of 59.1% for CH₃OH.³³ X-ray absorption spectroscopy and DFT calculations revealed that single atom copper can lower the energy barrier for the conversion of HCOOH* to adsorbed CHO* intermediate due to its unsaturated electronic structure. Fig. 1e shows the TEM image of the 2D MXene with copper single atom.

2.5 Metal–organic frameworks and covalent organic frameworks

Metal–organic frameworks (MOF) composed of metal and organic ligands are a class of porous, crystalline materials and have been intensively studied in various fields due to their tunable structures, vast porosity, and large surface area.⁷⁹ Notably, ultrathin 2D MOF nanosheets with larger surface area, adjustable chemical compositions and electronic structures, and more uniform pores can provide numerous surface active sites and greatly promote ion and electron transfers, making them highly suitable materials for CO₂ conversion applications.^{80,81} Cui *et al.* systemically studied CO₂ reduction activity for a family of emerging two-dimensional metal–organic frameworks (MOFs) *via* density functional theory (DFT) calculations.⁸² They found that Mo-based MOF exhibits good activity of CO₂ activation, and the energy barrier in Mo-based MOF for CO₂ reduced into methane is only 0.27 eV.⁸² These corresponding energy inputs in CO₂ reduction are substantially lower than traditional pure metal catalysts and some single atom catalysts with noble metal. Recently, a layered, bimetallic 2D conjugated metal–organic framework (2D c-MOF) with copper-phthalocyanine as ligand (CuN₄) and zinc-bis(dihydroxy) complex (ZnO₄) as linkage (PcCu-O8-Zn) was created and exhibited a high CO selectivity of 88% with excellent durability.⁸³ This performance is attributed to the synergistic effects between highly efficient ZnO₄ catalysts and CuN₄ centers. This study reveals the feasibility and advantages of utilizing bimetallic MOF catalysts. Similar to MOFs, 2D covalent organic frameworks have also exhibited outstanding CO₂ reduction efficiency and selectivity. Lin *et al.* developed covalent organic frameworks (COFs) composed of cobalt porphyrin catalysts linked by organic struts and demonstrated the advantages of such materials in catalytic CO₂ reduction.⁸⁴ As a result of the π conjugation and π–π stacking in the COF, the charge-carrier mobility is significantly improved, resulting in a high FE_{CO} of 90% and a turnover number of up to 290 000 at –0.67 V *vs.* RHE, demonstrating a 26-time increase in activity compared with molecular cobalt complex.⁸⁴ In addition, other 2D electrocatalysts can also be derived from metal organic frameworks and used for CO₂ ECR applications. For instance, 2D bismuthine has been derived from a Bi-based metal organic framework and demonstrated a high faradaic efficiency of more than 96% for formate production with large current densities and stability.³⁴ Fig. 1f and g show the morphologies of such MOF-derived structures. This example has provided novel approaches to utilize MOF in the field of CO₂ ECR.

3 Regulating strategies for 2D electrocatalysts for CO₂ reduction reaction

Many advanced strategies have been introduced to improve the CO₂ electrocatalytic performance of two-dimensional materials

by increasing the number of active sites and enhancing the activities of active sites. Generally, common approaches include structure and phases engineering, defect engineering, and nano composites.^{85,86} The desired outcomes of such optimizations are (1) higher current density at low overpotentials towards high-value products, (2) greater selectivity in single products and inhibiting the competitor HER reaction, and (3) enhancing long-term durability in catalytic processes.³ This section will review some advanced strategies to optimize 2D solid catalysts for CO₂ electrocatalytic reductions.

3.1 Structure engineering

Surface functionalization in solid catalysts is one of the most powerful methods to directly influence catalytic properties *via* modulating the electric and absorption properties of material surfaces. A recent study by Cheng *et al.* used a surface nitrogen-injection approach to reach high activity and selectivity for CO₂ conversion into formate. The nitrogen abundant Sn nanosheets (N-Sn) showed a 5-times higher current density and a 2.45-times higher faradaic efficiency than that of pristine 2D Sn nanosheets, resulting in a high productivity of 1358 $\mu\text{mol h}^{-1} \text{cm}^{-2}$.²⁷ The nitrogen abundant Sn surface has a higher CO₂ adsorption capacity than pristine Sn surface, because of the coordination between Sn and N atoms. Additionally, the undercoordinated Sn and defect sites created a higher active surface area of 90.79 mF cm^{-2} in N-Sn, much higher than the 64.18 mF cm^{-2} found in pristine S nanosheets.²⁷ DFT calculations revealed that the valence electrons of S atoms were increased from 6 to 6.86, suggesting a more positively charged Sn^{δ+} in N-Sn samples that resulted in higher reaction activity and selectivity.²⁷

Surface defect engineering has been a hot research topic in heterogeneous catalysis. Current research suggests that the abundant oxygen vacancies in 2D transition metal oxides, such as Co₃O₄ and Bi₂O₃, can promote the CO₂ adsorption on catalyst surface and accelerate the desorption of products from surface, thus reducing the rate-limiting activation energy and improves the activity and durability in CO₂ reduction reaction.^{87,88} A recent study on oxygen defective Bi₂O₃ nanosheets revealed that oxygen vacancies can result in a negative adsorption energy of -0.3 eV on the catalyst surface and promote the generation of CO₂^{•-} radicals *via* a single-electron transfer process.⁸⁸ The oxygen defective Bi₂O₃ demonstrated localized electrons around the oxygen vacancies, thus enhancing the interactions with CO₂ on catalyst surface.⁸⁸ The impacts of oxygen vacancies on CO₂ ECR performance were also investigated with Co₃O₄ atomic layers. It was discovered that the oxygen(II) vacancy decreased the activation barrier by stabilizing the HCOO⁻* intermediates and promoting the proton transfer. Consequently, a high formate selectivity of 85% was achieved in vacancy rich 2D nanocatalyst.⁸⁷

Furthermore, rich cationic vacancies and sulfur vacancies can promote the formation of the intermediate in CO₂ reduction using 2D nanocatalysts, leading to improved selectivity, activity, and stability of CO₂ reduction. Peng *et al.* reported the enhanced CO₂ ECR to *n*-propanol using CuS 2D

catalysts with double sulfur vacancies.⁸⁹ It was discovered that the highly dense double sulfur vacancies in copper sulfide (100) planes can act as electrocatalytic centers to stabilize CO* and OCCO* dimers and promote the CO-OCCO coupling to yield C₃ products. This phenomenon was not observed in CuS with single or no sulfur vacancies. Consequently, a high faradaic efficiency of 15.4% for *n*-propanol was observed.⁸⁹ In addition, some other research showed that defect engineering in graphene is of vital importance for tuning the electronic structure and catalytic performance.^{90,91}

The morphologies of 2D materials can play an important role in determining the CO₂ ECR performance. Structure engineering can significantly enhance CO₂ ECR performance by modulating the number of exposed active sites and affecting the chemisorption. Using a halogen-assisted calcination approach, Shu *et al.* synthesized 2D nitrogen-doped carbon from ZIF-8 precursor with a large electrochemical surface area of 788 to 1413 m^2 and a high G-N ratio of 42.27 to 55.06%.⁹² When the surface area increased from 788 $\text{m}^2 \text{g}^{-1}$ in the N/C-1100 sample to 1413 $\text{m}^2 \text{g}^{-1}$ in the N/C-Cl-1100 sample, the resulted current density showed a pronounce improvement, and the CO selectivity increased from 91.2% at -0.7 V to 99.5% at -0.5V.⁹² Fig. 2a shows the FE of catalysts with different surface areas. Additionally, the highest G-N ratio was also found in the N/C-Cl-1100 sample, suggesting a high active sites exposure. It was proven in this study that the increased G-N exposure can significantly promote the CO₂ electrochemical conversion to CO. This significant enhancement demonstrated the effect of surface area and active site exposure on the catalytic performance in 2D materials.

Modulating the exposure of active edges sites is another effective approach to enhance CO₂ ECR performance. DFT calculations have suggested that edge sites are highly active catalytic sites for CO₂ ECR using nanoparticles, calling for tunable 2D structures with high edges-to-corner site ratios.⁹³ Layered 2D structures have demonstrated superior performance to their bulk counterparts in CO₂ ECR, due to the exposure of more edges and defects for catalytic reactions.^{85,94} Recently, Li *et al.* synthesized layered Sb nanosheets by cathodically exfoliating bulk Sb crystals and demonstrated enhanced CO₂ ECR to formate performance.⁹⁴ With the same layer thickness, smaller Sb nanosheets with an average lateral size of 170 nm showed a 3.6-times higher peak current (mass loading normalized) than showed in larger Sb nanosheets with an average lateral size of 485 nm at a potential of -0.96 V, resulting in a FE of 84% for formate at -1.06 V.⁹⁴ The morphology-dependent CO₂ ECR performance of nanoplates was also observed with triangular silver nanoplates. Liu and coworkers discovered that Ag nanoplates demonstrated higher faradaic efficiency and energy efficiency than their Ag nanoparticle counterparts with similar sizes.⁹³ This enhancement was attributed to the larger amount of edge sites and higher edge-to-corner ratios in nanosheets, as the adsorption of CO₂ and the formation of COOH* are more promoted on edge sites.⁹³

The exposure of different facets is also a critical parameter in CO₂ ECR that determines the CO₂ ECR efficiency and



Fig. 2 (a) FE of catalysts with different surface areas. Copyright 2021, Springer Nature; (b) free energy diagram of CO₂ ECR to CO on different NG surfaces. Copyright 2016, American Chemical Society; (c) free energy diagrams and (d) the calculated limiting potentials for CO₂RR and HER on composite and pristine graphene surfaces. Copyright 2018, John Wiley and Sons. All panels reproduced with permission.^{23,92,96}

selectivity. Surface reconstruction has become an effective approach to program the desired exposure of facets, thus enhancing the CO₂ ECR of 2D materials. A recent work demonstrated that the preferred exposure of Bi (1 $\bar{1}0$) facets causes an increased FE of 90% for formate at a high current density of 200 mA cm⁻², due to the lower reaction energy barrier (0.03 eV) for CO₂ ECR exhibited on the Bi (1 $\bar{1}0$) facet than the energy barrier (0.74 eV) on the Bi (121) facet.²⁶ Moreover, the stepped Bi (1 $\bar{1}0$) surface also showed a Gibbs free binding energy of -0.01 eV for HCOO*.²⁶ This value is very close to the thermodynamic minimum. The facet-dependent performance was also observed in other nanosheets. For Ag nanoplates, the Gibbs free energy (ΔG) for forming COOH* was found to be lower on Ag (100) than on Ag (111), making Ag (100) the preferred facet for CO₂ ECR using nanoplates.⁹⁵

In 2D materials, doping is an extremely effective approach to tune the catalytic performance. Unlike in bulk materials, where dopants are located within the materials, doped 2D materials have more dopants exposed on the surface that can affect the overall catalytic performance. Nitrogen doping is a successful approach to enhance the CO₂ ECR performance using 2D carbon materials. In a recent study, three types of N dopants on graphene were prepared and investigated by Wu *et al.*: graphitic N, pyridinic N, and pyrrolic N, where the lowest uphill free energy barrier of COOH* adsorption was

observed in samples with triple-pyridinic N dopants.⁹⁶ Fig. 2b shows the free energy diagram of CO₂ ECR to CO on different graphenes. It was found that the positively charge C atoms adjacent to N dopants are the active sites for CO₂ ECR. Additionally, DFT calculations suggested that N doping can significantly reduce the energy barrier for COOH* adsorption, and the direct protonation of COOH* requires significantly lower energy barrier (-0.5 eV) to form CO gas and H₂O than that needed for the Heyrovski step to form H₂ (-2.0 eV).⁹⁶ Thus, this doping approach can greatly promote the CO formation and inhibit the competing HER. Besides, through computational approaches, Liu and coworkers discovered that pyrrolic-nitrogen doped graphene showed a lowest theoretical over potential of 0.24 V for CO₂ ECR to HCOOH among all N-doped graphene.⁹⁷ This result suggests that pyrrolic-nitrogen doping is a promising approach for CO₂ ECR to formate. However, the role of active site in N-doped graphene materials is still controversial in CO₂ electrochemical reduction reactions. Chai and Guo found that the graphitic N is the effective active site in electrochemical CO₂ reduction and has a lower activation barrier (0.58 eV).⁷² Furthermore, Sreekanth *et al.* also studied CO₂ electrochemical reduction of the B-doped graphene in aqueous 0.1 M KHCO₃ solutions.⁹⁸ Their research results showed that boron dopants changed asymmetric charge and spin density distribution of pristine graphene, and

B/C atoms with positive spin density in graphene are important active sites.⁹⁸

3.2 Composite materials and alloys

Engineered composite materials with heterostructures can often achieve enhanced CO₂ ECR performance by promoting charge transfer and the coupling effects between different phases. The general regulation parameters mainly involve compositions and alloys. Loading metals/metal oxides on two-dimensional graphene is a very common method to improve the electrical conductivity of metal oxides and increase the active sites of graphene. Bi *et al.* reported that the surface immobilization of Ni on nitrogen-doped graphene (Ni²⁺@NG) demonstrated significantly improved CO₂ ECR performance with a high current density of 30 mA cm⁻², compared with that of pristine N-doped graphene (<10 mA cm⁻²).²³ The FE for CO also experienced a drastic increase from below 60% to 92% in Ni²⁺@NG.²³ By comparing the free energy barrier for the formation of *COOH, *CO, and *H, the authors found out that Ni²⁺@NG showed a lower free energy barrier for the formation of adsorbed *CO than N-doped graphene did. Besides, compared with its pristine counterpart, Ni²⁺@NG composite 2D catalysts exhibited a higher preference towards CO₂RR over HER, as suggested by the limiting potential calculations. Fig. 2c and 1d show free energy diagrams on different graphene surfaces and the calculated limiting potentials for CO₂RR and HER, respectively.

Loading isolated metal atoms on graphene is an emerging regulation strategy in heterogeneous catalysis.⁹⁹ These composite graphene materials have demonstrated superior CO₂RR performance due to their high electrical conductivity, abundant active sites, and maximum atom efficiency. This strategy is regarded as one of the most effective strategies for improving the activity, stability, and selectivity in electrochemical CO₂ reduction. Combined with the attributes of single-atom catalysts and 2D materials, some researchers also synthesized an atomically dispersed Fe on nitrogen-doped graphene (Fe/NG), where graphene oxide (GO) absorbed iron ions (a mixture of GO and FeCl₃) is annealed in an atmosphere for Ar/NH₃ at 700–800 °C.¹⁰⁰ Fe/NG catalyst had a large number of nitrogen-confined atomic Fe moieties on the surface as active sites, resulting in a high faradaic efficiency of ~80% for CO production at -0.60 V *versus* RHE.¹⁰⁰ Theoretical calculations and experimental results suggested that the isolated Fe–N₄ moiety in nitrogen-doped graphene as the activity center that enhanced CO₂ adsorption and improved CO₂ activation. Based on these works, a heterogeneous molecular Co electrocatalyst (NapCo@SNG) was prepared by immobilizing planar CoII-2,3-naphthalocyanine complexes (NapCo) on S/N/O heteroatoms doped graphene (SNG) through axial coordination.¹⁰¹ Sulfoxide dopants in NapCo@SNG can enhance the electron transfer between NapCo and doped graphene, causing an increase of ~3 times in the turnover frequency of Co sites when compared with carboxyl-NapCo moiety and a high faradaic efficiency for CO production of ~97% at -0.8 V *versus* RHE.¹⁰¹ In addition, Cu/N-doped graphene composite

materials (p-NG/Cu) were introduced by combining mono-disperse Cu nanoparticles and pyridinic-N rich graphene (p-NG).¹⁰² The p-NG catalyst demonstrated the reduction activity and selectivity to ethylene compared with Cu and p-NG, with a faradaic efficiency of ~19% for C₂H₄ production and a 79% ethylene selectivity at ~0.9 V *versus* RHE.¹⁰² The interaction between p-NG and ~7 nm Cu NPs facilitated CO₂ reduction, hydrogenation, and C–C coupling in the formation process of C₂H₄.

Monolayer ternary transition metal dichalcogenide (TMD) alloys have demonstrated enhanced CO₂ ECR performance compared with binary TMDs. Xu *et al.* synthesized MoSeS alloy monolayers with shortened Mo–S and extended Mo–Se bonds that can tune the d-band electronic structure of Mo atom.⁶⁵ By engineering the band structure in this way, the density of states around the conduction band edge showed a drastic increase, resulting in faster electron transfers due to the lowered work function (4.24 eV) and decreased charge-transfer resistance.⁶⁵ Moreover, the overlap between the d-orbitals of Mo and the p-orbitals of Se atoms was reduced by the off-center charge density. DFT calculations revealed that the biased electron density around Mo atom in alloy monolayers can help stabilize COOH* intermediates and facilitate the CO desorption. Consequently, at an overpotential of -1.15 V *vs.* RHE, a high current density of 43 mA cm⁻² and a high faradaic efficiency of 45.2% for CO were observed in ternary TMD alloy monolayers.⁶⁵ Nevertheless, the low FE still needs to be improved, and the overpotential needs to be lowered to inhibit the competing HER.

4 Electrolyzers for electrocatalytic CO₂ reductions

Electrolyzers also play a decisive role in determining the CO₂ reduction performance of 2D materials. So far, various electrolyzer configurations have been developed and utilized in CO₂ ECR applications. The H-type cell, a simple configuration that employs an ion-exchange membrane to separate anodic and cathodic reactions, has been a popular electrolyzer configuration for 2D catalyst development and screening.^{26,89,103,104} Catalysts can be deposited on carbon paper to form the cathode for CO₂ ECR, and the reaction products are collected directly from the cells to evaluate CO₂ ECR performance. Nevertheless, the use of H-cells is greatly limited by their low current densities (<100 mA cm⁻²) because of the low solubility of CO₂ in liquid electrolytes.¹⁰⁵ In order to solve this issue, liquid-phase flow cells have been developed to directly use CO₂ gas as the CO₂ source in CO₂ ECR. Gas diffusion electrodes are applied in this configuration to allow for gas-phase CO₂ diffusions through the porous electrode and reactions taking place on the catalyst–electrolyte interface.¹⁰⁶ Ion exchange membranes, *i.e.*, anion exchange membrane, cation exchange membrane, and bipolar membrane, are utilized to separate the anodic and cathodic chambers. Improved CO₂ ECR stabilities, current densities, and faradaic efficiencies have been widely

reported with the use of flow cells.^{105,107} In addition, modified H-cell and flow cell configurations are the most common reaction cells for *in situ* and *operando* investigations.^{27,28,108,109} On the other hand, the membrane electrode assembly (MEA) is a novel gas-phase electrolyzer that uses a solid polymer electrolyte to separate the anode and cathode. The catalyst on the gas diffusion electrode is directly pressed against the membrane to form the “zero-gap” configuration that reduces the ohmic losses and maximizes the energetic efficiency.¹¹⁰ Consequently, high current densities ($>1 \text{ A cm}^{-2}$) and faradaic efficiencies have been reported using MEA.^{111–113} However, the operation of gas-phase reactors requires humidity that can be provided either by liquid electrolytes in the anode and/or humidified gas.¹⁰⁵ The separation of CO_2 ECR products from liquid electrolytes can be eliminated using all-solid-state electrolyzers, which adopt solid electrolytes to facilitate the continuous production of pure liquid products.¹¹⁴ Remarkably, highly concentrated ($\sim 98 \text{ wt\%}$) formic acid and acetic acid have been produced with this solid-state electrolyzers.^{115,116} It can be projected that the developments and optimizations of novel electrolyzers can enhance the CO_2 ECR performance of 2D materials and facilitate *in situ* and *operando* studies to better understand reaction mechanisms.

5 Advanced *in situ* and *operando* characterization technology

During CO_2 reduction reactions, catalysts are in metastable states and subjected to constant changes in their chemical and structural properties in a complex reaction environment. Such changes include surface reconstruction, catalysts deactivation, agglomeration, phase segregation, metal reconstruction, and vacancies formation.^{117–123} Many reaction intermediates are formed and consumed during the CO_2 RR process to yield various final products *via* different pathways.²⁸ On the

other hand, the changes in catalysts' chemical and structural properties can affect their catalytic performance, determining their effectivity and selectivity.¹²⁴ As illustrated in Fig. 3,²⁹ the elementary transferring of electrons and protons, the adsorption and desorption of intermediates and products, the mass transfer of molecules in the reaction environment, and other processes in catalytic reactions, all take place in different time and length scales. It is critical to study all such processes in a time resolved manner to gain insightful information about CO_2 catalysis. Nowadays, thanks to the developments of highly advanced *in situ* and *operando* characterization methods, fundamental studies of CO_2 RR at various time and length scales have become feasible. *In situ* characterizations refer to the characterizations of catalysts under conditions relevant to catalytic operation conditions. Such characterizations might not be simultaneous to reactions. However, *operando* characterizations are simultaneous characterizations of working catalysts during real catalytic reactions.^{125,126} In this section, state-of-the-art *in situ* and *operando* spectroscopy and microscopy methods are reviewed, and future research trends are proposed with the goal of achieving better understandings the mechanism and design guidelines of 2D materials in CO_2 RR. We summarized the spatial and temporal resolutions of these advanced characterization techniques in Table 1.

5.1 *In situ* scanning probe microscopy

5.1.1 *In situ* scanning electrochemical microscopy (SECM). Scanning electrochemical microscopy is a type of scanning probe microscopy that utilizes an ultramicroelectrode (UME) or a nanoelectrode (NE) to probe the local current in the vicinity of a spot of interest, thus detecting the chemical species in this area.¹²⁷ This technique has been employed in CO_2 RR studies to investigate the formation of different products and intermediates.^{128–132} Typically, the SECM system consists of an electrochemical cell with two working electrodes: working electrode 1 being the catalyst material, working electrode 2 being



Fig. 3 Time scale of various processes in catalytic reactions. All panels reproduced with permission.²⁹ Copyright 2017, John Wiley and Sons.



Fig. 5 (a) Schematic of *in situ* EC-STM setup. Copyright 2018, John Wiley and Sons; (b) schematic of *in situ* EC-AFM; (c) morphology of Cu surface at different potentials characterized by *in situ* AFM. Copyright 2020, John Wiley and Sons. All panels reproduced with permission.^{139,143}

to monitor the construction of Cu (100) and Cu (511) facets at different potentials and prepared oxide-free Cu (511) structures.¹⁴⁰ With additional quartz-crystal nanobalance and differential electrochemical mass spectrometry characterizations, the authors investigated the CO adsorption behaviors and proved that effective CO dimerization can only occur after a certain degree of CO coverage on the catalyst surface, leading to the formation of $\text{CH}_3\text{CH}_2\text{OH}$.¹⁴⁰ The same group of researchers also reported the cathodic regeneration of Cu (100) from air-oxidated surfaces under a typical CO_2RR condition.¹⁴¹ Additionally, they revealed the evolution of polycrystalline Cu to Cu (111) after 30 min of experiment in a 0.1 M KOH solution, at a CO_2RR potential of -0.9V .¹⁴¹ Interestingly, after another 30 min, Cu (111) further transformed into Cu (100). By adopting such time resolved monitoring of Cu surface, researchers hope to construct the “structure–composition–reactivity correlations” of heterogeneous electrocatalysts.¹⁴²

5.1.3 *In situ* atomic force microscopy (AFM).

Electrochemical Atomic Force Microscopy (EC-AFM) is another powerful scanning probe microscopy that is capable of investigating the catalyst surface morphology and surface potentials in *in situ* CO_2 ECR studies. A typical EC-AFM setup consists of a three-electrode electrochemical cell and an AFM with its tip submerging in the liquid electrolyte. For morphological studies, the amplitude modulation mode (AM-AFM) can be used to scan the surface of catalysts during reactions.¹⁴³ Fig. 5b illustrates the mechanism of this *in situ* measurement technique. On the other hand, the measurement of catalyst surface potentials involves a conductive AFM tip that directly contacts the sample and measures the catalyst surface poten-

tials.¹⁴⁴ Using *in situ* EC-AFM, Grosse and coworkers observed the change in Cu nanoparticles during a CO_2 ECR process. During the experiment, they observed drastic structural changes in Cu nanocubes. The loss in sharp corners and edges was observed after 1 min of CO_2 ECR process, and a 10% decrease in the particle size was probed after 3 hours.¹⁴⁵ Later, a correlation between such changes and the CO_2 ECR FE and selectivity was constructed. Smaller 220 nm-cubes were found to be more prone to morphological changes, leading to a drastic decrease in EF and an increase in selectivity towards C_1 products.¹⁴⁵ Recently, Simon *et al.* investigated the potential-dependent changes in the morphology of Cu catalyst in CO_2 ECR.¹⁴³ They found that at a low potential of -0.5V (RHE), the Cu (100) electrode surface exhibited round, smooth island and terrace structures. At -1V (RHE), rectangular islands with step edges replaced round islands, indicating the formation of low-coordination sites. However, at -1.1V (RHE), the size of rectangular structures reduced due to cathodic corrosion. Moreover, with the help of atomically resolved *in situ* imaging, the morphological change due to surface absorption was also detected with a $p(2 \times 2)$ superstructure present at -0.5V (RHE) and a (1×1) surface termination at -0.1V (RHE).¹⁴³ Fig. 5c shows the morphology of Cu surface at these two potentials. This work exhibited the outstanding *in situ* measuring capability of EC-AFM with an atomic resolution.

5.2 *In situ* and operando microscopy and imaging

5.2.1 *In situ* and operando transmission electron microscopy (TEM).

In situ and *operando* TEM can provide real time observations of the morphological properties of solid,

gaseous, and liquid reactants and catalysts with high spatial resolutions. It is a versatile technique that can be employed to study both liquid phase and gas phase reaction environment at atmospheric pressure, which can reflect the realistic working conditions for most of catalytic reactions.^{146–148} A high vacuum environment is usually maintained in the traditional TEM. However, *in situ* and *operando* studies require a real reaction environment around the sample. To meet this requirement, two methods have been developed to introduce gas and/or liquid to the reaction environment: (1). modified TEM. With Environmental Transmission Microscopy (ETEM), gas can be introduced to the sample using a differential pumping mechanism, which employs a set of differential apertures to control different pumping and pressure conditions inside the TEM. Consequently, a higher pressure at the sample area (~a few millibars) can be achieved with the introduction of reactant gas, while the rest of the ETEM column is still maintained at a high vacuum.^{147,149,150} (2). Modified sample holder. In this approach, the sample and reaction environment are kept within a sealed micro device, which is inserted to the TEM column to encapsulate the reaction environment while maintaining a high vacuum in the TEM column.^{146–148} Electron-transparent windows, usually made of silicon nitride membranes, are used above and under the sample for electrons to transmit.¹⁵¹ Notably, advanced micro devices can be fabricated by integrating a micro electrochemistry cell and a TEM sample holder, enabling electrochemical reactions to take place and be characterized inside the *in situ* TEM.^{146,152} This application is particularly helpful in electrochemical

CO₂RR studies using 2D materials. Fig. 6a illustrates the mechanisms of crystal growth in an electrochemical cell using *in situ* and *operando* TEM characterization approaches.¹⁵²

Benefited from *operando* TEM technology, our group studied the facet-dependent CO₂ catalytic reduction behaviors of Cu₂O nanoparticles and discovered that the (110) facet of Cu₂O nanoparticles is catalytically active in CO₂RR, while the (100) facet is not active.¹⁴⁸ The *operando* TEM was used to identify nanoparticles and facets of interest and guided the use of focused electron beam and nano-focused synchrotron X-ray beams to desired facets for localized spectroscopic investigations using electron energy loss spectroscopy (EELS) and X ray absorption spectroscopy (XAS).¹⁴⁸ In a recent publication by Wang *et al.*, a micro electrochemical cell (“*in situ* TEM E-chip cell”) was employed to investigate the real-time morphological changes in 2D CuO nanosheets in a CO₂ ECR environment.¹⁵³ They discovered the potential- and time-dependent conversions of 2D CuO nanosheets into nanoscale fragments, which consecutively agglomerate and form Cu dendrites. The formation of undercoordinated Cu sites was found to promote the CO₂ electrochemical conversions into C₂₊ products. This *operando* TEM successfully revealed the operation condition–morphology–performance relationship of CO₂ ECR using CuO nanosheets as the catalyst. Additionally, Vavra and coworkers recently monitored the dissolution/redeposition of Cu nanocatalyst in CO₂ ECR.¹⁵⁴ They found that, at the open circuit voltage, the oxidation and the dissolution of Cu nanoparticles took place at the same time, yielding Cu₂O and dissolved Cu ions in the initial electrolyte. Later, at higher catho-



Fig. 6 (a) The schematic of a micro electrochemical cell for *in situ* TEM. Copyright 2017, IOP Publishing, Inc.; (b) the change in primary (I) and secondary (II) particle sizes as experiment progresses; (c) the growth of secondary particles, and (d) the shrinkage of primary particles. Copyright 2020, John Wiley and Sons. All panels reproduced with permission.^{152,154}

dic potentials, the dissolved Cu^+ and Cu^{2+} ions are reduced back to solid Cu phase, generating larger secondary Cu nanoparticles. It was also proposed that the Cu surface oxide, Cu_2O , participated in the reconstruction process of Cu nanocatalyst as the CO_2 ECR experiment progressed.¹⁵⁴ With the help of time resolved monitoring at the nanoscale, the catalyst phase change mechanisms were investigated. Fig. 6b depicts the change in primary and secondary particles. Fig. 6c and d show the *in situ* observation of the morphological change in primary and secondary particles.

Nevertheless, there are still some limitations in such *in situ* and *operando* TEM characterizations. For example, it is still challenging to get high-quality 3D reconstructed images using *in situ* and *operando* TEM. One reason is due to the limited tilting angles inside TEM if using electrochemical cell. The gas bubbles generated in CO_2 RR might affect the local reaction environment, and this impact is much amplified in the micro reaction system where only a small amount of electrolyte is present.¹²⁴ Additionally, the CO_2 saturation in electrolyte results in deteriorated resolution and lowered contrast, due to inelastic electron scattering in liquids.¹⁵⁴ It was also reported in the same publication that a high electron dose can damage the catalysts within a short amount of time. Hence, the electron dose needs to be fine-tuned to avoid damaging the catalysts while achieving an optimal resolution.

5.2.2 *In situ* and *operando* scanning transmission X-ray microscopy (STXM). Synchrotron-based *in situ* Scanning Transmission X-ray Microscopy (STXM) is a powerful tool to characterize active electrochemical cells by providing high spatial and temporal resolution images and spectra with excellent chemical, electronic, and bond orientation contrasts.¹⁵⁵ In

STXM, synchrotron and undulator-generated monochromatic X-rays first pass through a zone plate that focuses X-rays onto the sample plane, and the sample is raster-scanned by the focused X-rays. Then, the transmitted X-rays are measured by detectors to construct images. With an additional fluorescence detector installed, Scanning X-Ray Fluorescence Microscopy (SXFM) can be combined with STXM to detect the fluorescent photons and reveal the distributions of elements on samples with unparalleled sensitivity.¹⁵⁶ Also, the X-ray absorption properties of materials can provide chemical bonding information and elemental composition information with high chemical contrast.¹⁵⁷ Fig. 7a illustrates the setup of STXM and SXFM. Advanced zone plate lenses and setups have enabled STXM to reach a spatial resolution of 7 nm and a temporal scanning resolution of 20–30 ps.^{158,159} The spectral resolution of soft X-rays at some synchrotron facilities can be as high as 0.1 eV.¹⁵⁷ By adopting an electrochemical nanoreactor, *in situ* and *operando* STXM characterizations of working electrocatalysts can be conducted at atmospheric pressure and up to 350 °C.¹⁶⁰ STXM images and spectra can probe critical information about catalytic reactions, including catalysts morphologies, the changes of catalyst phases and chemical states, as well as the properties and locations of catalytic reactants.^{160,161} In a recent study conducted by our group, *operando* SXFM was employed to conduct imaging and nanospectroscopic investigations of Cu_2O nanocatalysts for CO_2 reductions. The location and morphology of a single Cu_2O nanoparticle of interest were determined by low-resolution and high-resolution SXFM images, respectively. The corner of the particle was revealed to be the (110) facet, while the edge is the (100) facet. By directing the nanofocused X-ray beam paral-



Fig. 7 (a) STXM and SXFM setup; Copyright 2014, National Academy of Sciences; (b) SXFM image of a Cu_2O nanoparticle; (c) scanning X-ray nanospectroscopy of the (110) facet; Copyright 2019, Springer Nature; (d) to (f) chemical contour images of the catalyst under different treatments; Copyright 2008, Springer Nature. All panels reproduced with permission.^{148,156,160}

l to different facets, the facet-dependent X-ray fluorescence spectra were obtained to evaluate the active sites during CO₂ reduction reactions. The shift observed in the spectra generated on the (110) facet under reaction conditions revealed that the (110) facet is active for CO₂ reduction.¹⁴⁸ Fig. 7b and c are the high-resolution image of a Cu₂O nanoparticle with facets identified and the scanning X-ray nanospectroscopy of the (110) facet under different conditions. Our group also utilized an electrochemical liquid cell with STXM and XAS to characterize the non-dissolution nature of electrodeposited species on electrodes. It was observed that the non-dissolvable Mg thin film during discharging was hexacoordinated Mg compounds.¹⁴⁶ This example is particularly relevant to 2D materials for CO₂ ECR applications, as similar *in situ* and *operando* techniques can be applied to investigate the surface reconstruction phenomenon of 2D catalysts during reactions. A study by Smith *et al.* demonstrated the capability of STXM to examine the morphological and compositional properties of a complex iron-based Fisher–Tropsch catalyst, as well as the properties and locations of the produced carbon species under working conditions. Chemical contour maps were constructed by processing STXM data files with an aXis2000 software package and revealed the compositional changes of catalysts during reactions, as shown in Fig. 7d–f: f. before treatment; e. after 2 h in H₂ at 350 °C; c. after 4 h in synthesis gas at 250 °C.¹⁶⁰

5.2.3 Coherent X-ray diffraction imaging. The emergence of Coherent X-ray diffraction imaging (CXDI) has offered a highly advanced *in situ* and *operando* characterization approach to characterize structural properties of nanomaterials. So far, CXDI has demonstrated its ability to reveal 3D material morphologies, strains, and defects, as well as to probe surface adsorptions on catalysts at a spatial resolution of less than 10 nm.^{162–168} In the CXDI process, an advanced X-ray source generates a spatially coherent X-ray beam, which then interacts with the sample and creates a far-field diffraction pattern. To cover the whole sample, the coherence volume of the beam should be larger than the sample size. Then, with an “oversampling” technique, a computational algorithm can retrieve the phase information from the diffraction. Traditionally, to obtain 3D diffraction patterns, the sample needs to be rotated to change the sample’s angle to the beam. Afterwards, 3D images with material structural information are constructed from the obtained 3D diffraction data by a computer algorithm. So far, many optimization methods have been developed to enhance the performance of CXDI to meet the requirements for faster and simpler characterizations, including “chrono CDI” method to reduce the oversampling requirement,¹⁶⁸ and the “variable-wavelength” CDI method to construct 3D diffraction without rotating the sample.¹⁶⁶ This technology can bring unique opportunities for the studies of 2D materials for CO₂ ECR. For instance, strain engineering has been recently demonstrated to be an efficient strategy for enhancing the CO₂ ECR activity and selectivity, but the quantification of strain is still challenging.¹⁶⁹ By using CDI, Fernández *et al.* *in situ* monitored the strain change of single

Pt nanoparticles upon exposure to O₂ and CO/O₂ mixtures.¹⁷⁰ An outstanding strain sensitivity of 10^{−4} was achieved to monitor the strain variation of the model catalyst with a 1-minute time resolution. They observed a tensile expansion of up to 0.09% in the nanoparticles when exposed to a O₂ environment and attributed this strain change to the formation of a platinum oxide that resulted in possible overheating of the sample and/or surface stress and lattice mismatch.¹⁷⁰ However, when a reducing environment was introduced by having an excessive amount of CO in the CO/O₂ gas mixture, compression started to occur, reducing the strain to less than 0.03%. Another *in situ* investigation on Pt nanoparticles conducted by Abuin and coworkers also examined the structural changes due to CO adsorption.¹⁶⁴ For the near surface region of the nanoparticle, a −0.01% strain was observed under Ar flow. However, in a mixed Ar/CO flow environment, the strain in the near surface region changed to +0.12%, suggesting that CO adsorption weakened the bonds among surface Pt atoms and thus generated an outward relaxation of these atoms. Consequently, new high index facets were created. Meanwhile, the bulk region exhibited a compressive strain of −0.15% under Ar flow and a decreased strain of −0.02% under AR/CO flow, implying CO-adsorption induced lattice expansion in the bulk region.¹⁶⁴ Fig. 8a illustrates the experiment setup and mechanism of CXDI characterization. Fig. 8b shows 3D CXDI imaging of Pt NP under (a) Ar flow and (b) mixed Ar/CO flow (c) probability density of the strain field in near surface region and (d) in the bulk.¹⁶⁴ In addition, they also found a good match between nanoparticle morphologies measured by *ex situ* AFM and *in situ* CXDI. As CO is an important product and intermediate in CO₂ ECR, such *in situ* studies using CXDI offered new aspects to be considered when studying and designing 2D materials for CO₂ ECR. Nevertheless, it must be pointed out that CXDI studies require advanced synchrotrons, which are mainly available at the 34-ID-C beamline at the APS, the EBS ID01 beamline at the ESRF, and a few others in Europe.

5.3 *In situ* and *operando* X-ray spectroscopy

5.3.1 *In situ* hard X-ray absorption spectroscopy. X-ray absorption spectroscopy (XAS) has been intensively used in catalytic reactions, including CO₂ ECR. In this characterization, X-ray beams generated by a light source interact with the sample and are partially absorbed. By measuring and plotting the absorption of X-rays as a function of photon energy, XAS can probe the dynamic changes in electronic structures of catalysts and thus be used for the studies of reaction mechanisms in CO₂ ECR processes. *In situ* and *operando* spectroscopy that probes the dynamic changes in a catalytic environment requires high intensity photons with a broad range of energies to construct highly resolved X-ray spectra. Usually, around 10¹⁰ photons per eV are needed, and this requirement can only be met by the use of synchrotrons to date.¹⁷²

Hard X-rays with photon energies higher than 5 keV have strong penetration power. Consequently, hard X-ray absorption spectroscopy is highly desired *in situ* and *operando* studies that



Fig. 8 (a) The schematic of CXDI imaging. (b) 3D CXDI imaging of Pt NP under different conditions and the strain field in near surface region and in the bulk. Copyright 2019, American Chemical Society; (c) *in situ* XAS electrochemical cell. Copyright 2017, Springer Nature; (d) time-resolved XAS spectrum. Copyright 2021, Springer Nature. All panels reproduced with permission.^{153,164,171}

involve complex reaction environments with the presence of liquid, gas, and solid. In XAS, the absorption spectrum ranges immediately before and after the absorption edge is defined as the X-ray absorption near edge structure (XANES), which can characterize the density of states, thus revealing the valence states (at the absorption edge) and orbital hybridization due to coordination (pre-edge).¹⁷³ The extended X-ray absorption fine structure (EXAFS) covers the energy region above the absorption edge and can characterize the local atomistic structure, *i.e.*, the neighboring atoms to the bound atoms and the interatomic distances between them. For *in situ* and *operando* CO₂ ECR studies, specially designed electrochemical cells with X-ray windows are required to provide a reaction environment that can be characterized by both electrochemical workstations and XAS. Such windows are usually made from Si₃N₄ to ensure no blockage for X-ray to pass through. Fig. 8c shows a design of this cell.¹⁷¹ *In situ* XAS has been utilized in the studies of 2D materials for CO₂ ECR as a powerful tool to probe the oxidation states and local coordination environment, thus determining the active sites during CO₂ electrochemical reduction reactions. For instance, a recent study on a bimetallic 2D conjugated metalorganic framework (2D c-MOF) demonstrated the synergistic effects between its CuN₄ ligand and ZnO₄ complex (ZnO₄) linkage with a high CO selectivity of 88%.¹⁰⁸ In this study, EXAFS spectrum showed an intensity decrease in Zn–O bond, suggesting interactions between ZnO₄ sites and reaction intermediates. With additional infrared absorption studies,

the authors were able to conclude that ZnO₄ sites are catalytic reactive sites. Another investigation on 2D bismuth metal-organic framework employed *in situ* XAS to probe the stability of Bi³⁺ during CO₂ ECR. XANES analysis proved that the Bi³⁺ oxidation state stayed unchanged at a high potential of ~ -0.9 V for 90 min, revealing a significantly enhanced Bi³⁺ stability in the prepared MOF structure than in other 2D materials.¹⁰⁴ Moreover, Wang and coworkers recently studied the highly selective CO₂ ECR reaction mechanism of 2D CuO nanosheet.¹⁵³ By using *in situ* XAS, they discovered that the CuO nanosheet gradually and directly converted into metallic Cu under a potential of -0.84 V (RHE) with no Cu₂O intermediates detected.¹⁵³ Fig. 8d shows the time-resolved information on the catalyst under a CO₂ ECR condition. EXAFS also showed that the coordination number of the produced Cu⁰ is smaller than that in bulk fcc Cu⁰, suggesting the existence of lattice defects associated with undercoordinated Cu atoms, which attributed to the enhanced efficiency and selectivity towards C₂₊ products. These in-depth mechanism studies demonstrated the abilities and advantages of *in situ* and *operando* XAS employed in CO₂ ECR studies using 2D materials.

5.3.2 *In situ* soft X-ray absorption spectroscopy. Soft X-rays absorption (sXAS) spectrum utilizes soft X-rays with energies below 5 keV and is a powerful tool to probe the structural properties of all the materials in a catalytic environment. For instance, the K-edge spectrum of the light elements such as carbon, nitrogen, and oxygen can be analyzed to study the

structural properties of 2D carbon-based catalysts and electrolytes.^{103,174} Moreover, the L/M-edge spectrum of transition metals can be more sensitive to oxidation states than K-edge spectrum.¹⁷² Additionally, soft XAS is very sensitive to the surface of solid materials and can reveal critical surface structure information, such as orbital hybridization, spin state, oxidation state, and adsorption species, *etc.* Heterogeneous catalysis mainly happens on the surface of solid catalysts, so the *operando* characterization of soft XAS is of vital importance to investigate the catalytic mechanism at atomic scales. Recently, many new discoveries and progresses made in CO₂ ECR investigations were benefited from *in situ* sXAS studies. Notably, *in situ* sXAS for the Cu L₃-edge demonstrated that the ratio and electronic state of copper species exhibited dynamic changes under different applied potentials in CO₂ ECR.¹⁷⁵ Time-resolved sXAS discovered that Cu²⁺ in Cu₂(OH)₃Cl can be easily converted into Cu⁺ species within 5 min of reaction, whereas the transition from Cu⁺ to Cu⁰ would take much longer time. This result suggested that the reduction of copper is hindered by oxychloride sol-gel, thus stabilizing Cu⁺ species at more negative potentials.¹⁷⁵ Moreover, *in situ* sXAS has also been employed to probe the stabilization of other novel CO₂ catalysts, such as La₂CuO₄ perovskite nanocatalyst,¹⁷⁶ and demonstrated its great potential in the development of advanced CO₂ ECR catalysts. However, it must be pointed out that due to lower photon energies, the operation of soft XAS usually requires an ultra-high vacuum. Hence, the *in situ* and *operando* electrochemical cell needs to be well designed.

5.3.3 *In situ* X-ray photoelectron spectroscopy. X-ray photoelectron spectroscopy (XPS) is another powerful surface analysis tool that can probe the elemental compositions, chemical state, and electronic structures of the material surface. The operation of XPS is based on the photoelectric effect: the irradiation of X-rays can excite electrons and result in photoelectron emissions to the vacuum. By measuring the kinetic energy of emitted electrons, the bonding energy of these electrons can be quantified to identify the surface elements and their chemical states. However, due to the short mean free path of the emitted photoelectrons, XPS is mainly used to probe the surface properties within a few nanometers of depth, and the operation of XPS usually requires ultra-high vacuum conditions (UHV), which limits the used of XPS in *in situ* CO₂ ECR studies.²⁸ A few approaches have been proposed to counter this issue, including the use of membrane-sealed “environmental cells” and highly advanced near-ambient pressure XPS (NAP-XPS) that utilizes the differential pumping mechanism.^{177–179} Similar to the microcells used in *in situ* TEM characterizations, the environmental cells can maintain atmospheric pressure within the cell while the whole cell is placed in a vacuum environment for XPS characterizations. The membranes, made of graphene oxide and other two-dimensional materials, are electron-transparent while strong enough to encapsulate all the reactants in the reaction environment.¹⁷⁹ In the differential pumping system, a set of pumping stages allows for an elevated pressure of a few milli-

bars in the sample area, from where the pressure gradually decreased to a high vacuum level in the detector area. Small apertures are used in such systems to maintain the pressure profile. Fig. 9a shows the mechanisms of these two approaches.¹⁷⁸

In situ XPS has been employed in some fundamental studies to scrutinize CO₂ ECR mechanisms on different catalysts. For example, Favaro *et al.* used *in situ* ambient pressure X-ray photoelectron spectroscopy to investigate how subsurface oxide below Cu surface affects the CO₂ activation process.¹⁸⁰ *In situ* XPS examined the binding energy shifts of C 1s and O 1s photoelectron peaks on different surfaces. It was found that the presence of an optimal amount of subsurface oxide can form a mixed Cu⁺ and Cu⁰ structure that stabilized the bent CO₂, thus promoting the CO₂ activation.¹⁸⁰ Similarly, Kim and coworkers used *in situ* XPS to directly monitor the formation of dissociated CO* from chemisorbed CO₂ on Rh surfaces in a time-resolved manner.¹⁸¹ Notably, this experiment was conducted under a 0.1 millibars environment. Fig. 9b shows the time-resolved monitoring of the formation of the dissociated CO adsorption on Rh surface.¹⁸¹ Note that all the above examples used synchrotron radiation sources for the *in situ* ambient pressure XPS characterization, especially with the presence of liquid in the system. Synchrotron light sources can modulate the photon energy and thus characterize the samples in different depths, enabling both depth-resolved and time-resolved characterizations.¹⁸²

5.4 Free electron laser (FEL) and “pump/probe” technique

Free Electron Lasers (FEL) can provide extremely coherent laser beams with ultrashort pulse duration (femtoseconds), ultrahigh brilliance, and widely tunable wavelengths ranging from microwave frequencies to hard X-rays.^{184,185} They can generate much brighter electromagnetic radiation than other sources available nowadays.¹⁸⁶ With such unique advantages, FELs can be used for both imaging and spectroscopy, offering ultrafast time-resolved investigations of material properties and molecular dynamics with a temporal resolution in the scale of femtoseconds. In X-ray Free Electron Lasers (XFEL), a high energy relativistic electron beam is first generated by an electron source and then passes through a periodically alternating magnetic field, which is provided by the undulator/wiggler and forces the electrons to move along a sinusoidal path. The undulating motions of electrons then emit highly coherent radiations with wavelengths that can be modulated by the electron energy and/or the magnetic field strength. Fig. 9c illustrates the configuration of the FEL system.¹⁸³ There are only a few FEL facilities in the world nowadays, located in Asia, Europe, and the U.S.A. Notably, the European XFEL, located in Germany, has a high beam energy of 8.5–17.5 GeV and a short pulse duration of 3–150 fs. The photon energy is ranged from 240 eV to 25 000 eV, making it a powerful tool for probing various information mentioned above *in situ* catalytic investigations.

These unique advantages of FEL enable it to be used in the ultrafast pump/probe technique. In this technique, first, an



Fig. 9 (a) Two NAP-XPS systems: differential pumping system and environmental cell system. Copyright 2018, John Wiley and Sons. (b) Time resolved NAP-XPS showing the formation of dissociated CO absorption. Copyright 2020, Springer Nature; (c) configuration of the Free Electron Laser (FEL) system. Copyright 2021, Springer Nature. All panels reproduced with permission.^{178,181,183}

ultrashort light pulse that lasts for less than 100 fs interacts with the material and initiates photochemical reactions, thus exciting the system to a non-equilibrium state. This step is called the “pump” step. Then the non-equilibrium state is probed by a following ultrashort light pulse that monitors the changes that occurred due to the excitation. The temporal intervals of these two pulses can be modulated to investigate the time-dependent dynamics of the reactions. Besides using coherent X ray as pump and probe, the ultrafast electron diffraction (UED) use the laser pump and electron beam probe to study the molecular structural dynamics.¹⁸⁷ Consequently, the field of “femtochemistry” was constructed to investigate the fundamental reaction mechanisms in chemical reactions.¹⁸⁸ The pump and probe technique have been adopted in CO₂RR studies to mainly investigate the carrier dynamics of photocatalysts, as such systems can be easily excited by pump pulses.^{189,190} However, many other reactants and catalysts in other catalytic systems cannot be excited by light pulses. Fortunately, hopes for probing electrocatalytic reactions have been brought by some pioneering studies. Ultrafast XAS,

which can be realized by the pump/probe spectroscopy, has been utilized in the *in situ* monitoring of the sub-femtosecond motions of valence electrons. The monitoring of such motions can help reveal the fundamental reaction mechanism of CO₂ ECR using 2D materials.

5.5 Other *in situ* and *operando* spectrometry

5.5.1 *In situ* and *operando* mass spectrometry.

Mass spectrometry (MS) is an efficient, sensitive, and accurate method to probe the catalytic intermediates and products during CO₂ ECR, offering deep understandings of the reaction dynamics and mechanisms. Unlike liquid chromatography and gas chromatography, which require minutes of time for analytes to flow through columns, MS measures the mass-to-charge ratios of ionized samples with a temporal resolution on the scale of a second.¹⁹¹ Hence, online electrochemical MS techniques, especially Differential Electrochemical Mass Spectrometry (DEMS), have been adopted in *in situ* studies to directly probe reaction products and, therefore, the local reaction environment on the electrode–electrolyte interfaces in CO₂ ECR.¹⁹²

DEMS is constructed with a three-electrode electrochemistry cell, a porous PTFE membrane that lets the gaseous products flow through, and a vacuum system that connects to the MS. DEMS can help probe the potential-dependent activity and selectivity of catalysts in a time-resolved manner.¹⁹³ Remarkably, online MS can be combined with other *operando* characterization tools, such as TEM, to construct the structure–activity relationships of various catalysts.^{124,194} Using this technique, Clark *et al.* demonstrated the direct observation of ten different reaction species in CO₂ ECR with Ag and Cu catalysts.¹⁹² Through the direct sample collection at the electrode–electrolyte interface, *in situ* DEMS probed larger amounts of aldehydes than that characterized by *ex situ* methods in bulk samples.¹⁹² This demonstrates better products collection efficiency by using *in situ* methods rather than traditional *ex situ* methods. This should be considered in the experimental design and when quantifying the efficiency and selectivity of the catalyst. Additionally, there were 2–4 times more aldehydes than the corresponding alcohols.¹⁹² Such results suggested that these aldehydes are reduced to corresponding alcohols prior to entering into the bulk electrolytes in CO₂ ECR. The potential-dependent product distribution was also characterized. Even though *in situ* MS has been utilized in many studies, future optimizations are still required to further improve the temporal and spatial resolutions.¹⁹⁵

5.5.2 *In situ* and *operando* infrared spectroscopy. *Operando* Infrared (IR) spectroscopy can monitor and analyze the reaction intermediates and products in real-time with relatively high temporal resolutions, greatly benefiting the investigations

on CO₂ ECR reaction mechanisms and catalytic sites.¹⁹⁶ The operation of IR is based on the specific absorptions of mid-infrared beams by molecules where the characteristic vibrational modes of molecules match the beam energies. Four operation modes are available for *in situ* or *operando* IR characterizations in catalytic reactions, including external-reflectance (*i.e.*, infrared reflection absorption spectroscopy, IRAS), attenuated total reflection (ATR), diffuse reflectance FT-IR spectroscopy (DRIFTS), and transmission (TIR).¹⁹⁷ In electrocatalytic systems that involve liquid electrolytes, the high infrared absorption of water and ionic liquids can interfere with such IR characterizations. Therefore, it is critical to shorten the path length of IR beams in *in situ* IR measurements. This requirement has made ATR and IRAS particularly suitable for *in situ* IR characterizations for electrochemical reactions. In IRAS, the tip of the working electrode, on which the electrochemical reactions take place, is placed facing an infrared-transparent window that only 1–10 μm of electrolyte layers present between the electrode and window. The IR beams then transmit through the window for measurements. Hence, the interference caused by electrolytes can be minimized.¹⁹⁶ ATR-IR adopts the Kretschmann configuration for attenuated total reflection. The working electrode is deposited as a thin film on an infrared-transparent prism, which allows the IR reach total internal reflection when the IR beam is placed at an angle of incidence larger than the critical angle. Thus, the IR beams do not pass through the liquid electrolyte and instead penetrate the liquid only by the depth of 1–2 μm through the evanescent wave to interact with adsorbed molecules on the surface of electrocatalysts.¹⁹⁸



Fig. 10 (a) Mechanism of *in situ* ATR. Copyright 2017, crown; (b) time-resolved SEIRA spectra of H_{ad} accumulation on CO-adsorbed Cu surface. Copyright 2016, American Chemical Society (c) mechanism of SERS. Copyright 2021, Royal Society of Chemistry; (d) *operando* Raman spectra under different applied potential. All panels reproduced with permission.^{198–201} Copyright 2021, 2021 Wiley and Sons.

Fig. 10a illustrates the mechanism of this advanced technique. By combining ATR and the Surface Enhanced Infrared Absorption Spectroscopy (SEIRAS), ATR-SEIRAS has demonstrated enhanced time resolution and sensitivity for *operando* studies, less interference from the bulk electrolyte, and the reduced IR attenuation and transport limitations.¹⁹⁶ The high IR absorption cross-section of C–O bonds makes IR very suitable for CO₂ ECR studies that produce formate, CO, as well as other common intermediates and products in CO₂RR.¹⁹⁸ SEIRAS has been widely utilized in CO₂ ECR studies for the investigations of fundamental reaction mechanisms and reaction species. For instance, SEIRAS demonstrated that the adsorbed hydrogen could partially displace adsorbed carbon monoxide on the Cu surface during CO₂RR, while the displacement can't occur in the reversed direction. Hence, it was proposed that hydrogen has a higher absorption strength than that of CO.¹⁹⁹

Fig. 10b shows the time-resolved SEIRA spectra of hydrogen and CO on H_{ad} pre-covered and CO pre-covered Cu surfaces, proving that hydrogen can displace adsorbed carbon monoxide.¹⁹⁹ Additionally, SEIRS also revealed that the surface C-bound and O-bound reaction intermediates can affect the overall CO₂RR reaction selectivity. Regardless of the powerful characterization ability, one drawback of *in situ* and *operando* IR in catalytic reactions is the limited access to low wavenumbers, which can help obtain information on the structural properties of surface oxides and hydroxides.¹⁹⁶ Moreover, the focused IR beam area is on the micron scale, which is relatively large for precisely probing nanomaterials. Hence, the measured peak only shows the averaged data within the whole beam area.

5.5.3 *In situ* and *operando* electrochemical Raman spectroscopy. Advanced *in situ* electrochemical Raman spectroscopy (RS), including surface-enhanced Raman scattering (SERS), shell-isolated nanoparticle-enhanced Raman spectroscopy (SHINERS), and tip-enhanced Raman spectroscopy (TERS), is a powerful and nondestructive tool to monitor the catalyst states and probe the reaction intermediates and products during catalytic reactions.²⁸ The operation of Raman spectroscopy is based on the Raman scattering, *i.e.*, the inelastic scattering of incoming photons due to the interactions between incoming monochromatic light beams and chemical bonds, resulting in wavelength shifts. Raman spectroscopy (RS) is complementary to IR. However, unlike the absorption-based IR, the scattering-based RS better suits the characterizations of an aqueous environment, as water doesn't induce strong Raman scattering.²⁰⁰ In *in situ* RS, the beam is introduced by a confocal system and interacts with the catalyst surface after passing through an observation window, through which the scattered signal is also collected using a reflective mirror.

Fig. 10c illustrates the schematic of this operation. As an enhanced technique, SERS operates based on the SERS enhancement effects, which are induced by the amplified local electric field as a result of the resonant excitations of surface plasmons and the non-resonant lightning rod effects due to

the concentrated field at enhancement sources, *i.e.*, surface features that maximize electromagnetic enhancement.¹⁹⁸ A significant enhancement in the Raman signal can be achieved in this way, making this technique highly desired in catalytic reactions to monitor the surface adsorbates and reaction species.²⁰⁰ Nevertheless, the dominant enhancement sources can also interfere with the characteristic signals from samples. Consequently, SHINERS was developed to isolate the enhancement sources (such as sharp tip/surfaces) to minimize the interference, making SERS applicable to all kinds of surfaces and morphologies.²⁰² On the other hand, by combining SERS and scanning probe microscopy tips, tip-enhanced Raman spectroscopy (TERS) was introduced to provide highly sensitive and spatial-resolved characterizations by involving a scanning probe microscopy tip. With a spatial resolution as high as 1–10 nm, TERS is extremely helpful in *in situ* characterizations of nanocatalysts.^{203,204} *In situ* Raman characterizations have been demonstrated with different 2D materials for CO₂ ECR and provided information on the reaction mechanisms and reactive sites. For example, in a recent study, bismuth phosphate (BiPO₄)-derived 2D Nanosheets were examined by *operando* Raman spectroscopy during CO₂ ECR at different potentials. As shown in Fig. 10d, the disappeared Raman peaks associated with the PO₄ group and the shifts of peaks assigned to Bi–O bond suggested the formation of Bi and BiO_x during CO₂ ECR.²⁰¹ The abundance of Bi–O on the surface was attributed to the enhanced CO₂ ECR performance. Many reaction intermediates, such as *CO and *HCOO, have been identified in *in situ* studies as key intermediates in CO₂ ECR.²⁰¹ Even though TERS is a highly advanced *in situ* characterization technique, many limitations, such as the short tip lifetime and the inability to characterize single-molecule level features, still need to be overcome in the future.²⁰³ Similar to IR, Raman spectroscopy is also limited by the large beam area with average data when characterizing nanomaterials, which make lack the heterogeneity information from samples.

6 Challenges and outlook

In recent years, 2D materials have been intensively investigated and utilized in CO₂ electrochemical reductions to value-added products. With the emergence of many innovative 2D catalysts, exceptional progress has been achieved *via* various approaches to enhance the activity, selectivity, and overall performance of such 2D catalysts, providing promising paradigms for the next generation CO₂ conversion systems. Nevertheless, the development of 2D catalysts for CO₂ ECR still faces many challenges, bringing promising opportunities in this field of study. Most of the current studies are only focused on the catalyst material aspects, such as the types of materials, morphologies, defects, and compositions. However, catalytic systems are often very complex and include numerous components that need to be integrated into the studies of 2D catalysts. Many parameters, including the reactor design, electrolytes, and local reaction

environment, can all determine the overall CO₂ ECR performance.

In CO₂ ECR, the local reaction environment can be constructed by both solid/liquid and solid/gas/liquid interfaces. With the adsorption, reaction, and desorption all taking place at these interfaces, these reactions are all governed by the synergistic effects of all the components in the local environment. It has been reported that the interactions between solutions and reactants can critically affect CO₂ performance. For instance, the types and sizes of both anions and cations in electrolytes have demonstrated governing effects on the selectivity of CO₂ ECR *via* affecting the adsorptions of reaction intermediates, altering the charge distributions on catalysts surface, and shifting the reaction free-energy landscape of Cu catalysts.^{205–207} Moreover, the electrical fields created in the Helmholtz layer and the use of ionic liquids have demonstrated various effects.^{208,209} Organic surfactants like cetyltrimethylammonium bromide have also enhanced CO₂ ECR performance by inhibiting the competing HER at the catalyst surface.²¹⁰ The interactions between liquid electrolytes and solid catalysts and the generated liquid–solid interfaces can all determine the CO₂ ECR performance. With their unique 2D morphologies and substantial dopant and vacancy exposures, 2D nanocatalysts might induce distinctive reaction behaviors in such a highly localized reaction environment. Such investigations on the local reaction environment are required to enhance the performance of 2D CO₂ catalysts further.

Researchers might also investigate additional novel enhancing strategies, such as catalyst modifiers and strain engineering, in the study of 2D materials. Recently, polymers have shown exceptional optimizing capabilities by modulating the heterogeneous catalytic surfaces and stabilizing metal nanocatalysts.^{211–213} Hydrophilic polymeric modifiers can enhance the selectivity for formic acid, while cationic hydrophobic modifiers can promote the formation of CO on Cu catalysts.²¹² Moreover, polymeric N-heterocyclic carbenes (NHC), such as polydentate and monodentate, have exhibited stabilizing effects on Au and Pd nanocatalysts due to the formation of metal–carbene bonds that inhibit the nanoclustering of nanoparticles.²¹¹ Such findings especially pertain to the use of 2D catalysts. The mechanisms and approaches proposed in novel optimization paradigms may also be applied to 2D nanocatalysts to achieve organic modifications at the atomic scale. Lattice strains have demonstrated the potential to alter the band structures of 2D materials and break the linear scaling relationship for CO₂ ECR, thus enhancing the selectivity and activity.^{214,215} However, it is still challenging to scrutinize the underlying mechanisms due to the difficulties of decoupling strain from other electronic effects in 2D materials with mismatched lattice.¹⁶⁹ More simulation and experimental approaches need to be combined to investigate such optimization behaviors. Additionally, applying strain on 2D catalysts with desired magnitude and distribution remains difficult to realize.

With the introduction of gas diffusion layers and polymer ion-exchange membranes in membrane electrode assembly

(MEA) electrolyzers, gas-phase electrolyzers have enhanced the current density and stability of CO₂ ECR.¹⁰⁵ Remarkably, some novel all-solid-state reactors that utilize polymer electrolytes can significantly facilitate the separation of liquid products in continuous productions.¹¹⁴ However, to date, most investigations on 2D catalysts have been conducted in liquid phase electrolyzers. The opportunities brought by such advanced reactors need to be seized in the study of 2D materials for CO₂ ECR. It must be pointed out that, in these solid-state reactors, the interactions between solid catalysts and solid electrolytes, the adsorption, desorption, and the mass transport of reaction species, as well as the local reaction environment are all different at the solid–solid interfaces from those at traditional liquid–solid interfaces. The behaviors and performance of 2D nanocatalysts in these advanced reaction environments and reactors need to be characterized and optimized to achieve superior CO₂ ECR efficiency and selectivity in future industrial-scale CO₂ ECR systems.

Due to the complexity of the catalytic reactions, many scientific problems cross orders of magnitude in times and length scales. Thus, we urge to use *in situ* and *operando* multimodal characterizations to get comprehensive, insightful information across different characterization platforms. Advanced *in situ* and *operando* multimodal characterizations of 2D nanocatalysts can help achieve fundamental understandings of CO₂ ECR mechanisms, especially when coupled with modern DFT calculations. Even though such characterizations often require advanced facilities such as synchrotron X-ray sources, electron microscopy, and free electron laser, the information corrected can reveal the underlying reaction active sites, pathways, intermediates, local environment, and morphological changes of 2D catalysts. Therefore, it is essential to actively employ *operando* multimodal characterizations in future studies to better design 2D nanocatalysts and beyond. More importantly, such *in situ* and *operando* characterizations conducted in flow cells and MEA electrolyzers remain extremely challenging. To develop 2D nanocatalysts used in industrial scale electrolyzers, specially designed apparatuses are highly demanded by *in situ* and *operando* studies.

Finally, to realize the utilization of 2D materials in real-life, industrial-scale CO₂ conversion systems, the mass production of 2D materials with high efficiencies, high selectivity and high stability must be achieved in the coming future. Yet, the production of 2D nanomaterials is still very challenging due to the aggregation and restacking caused by electrostatic and van der Waals interactions. Moreover, the productivity is still too low to achieve cost-effective productions and utilizations of 2D materials for CO₂ conversion. On the other hand, the feasibility of mass production should be kept in mind when designing next-generation 2D nanocatalysts.

Conflicts of interest

The authors do not have conflict of interests to declare.

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References

- C. D. Thomas, A. Cameron, R. E. Green, M. Bakkenes, L. J. Beaumont, Y. C. Collingham, B. F. Erasmus, M. F. De Siqueira, A. Grainger, L. Hannah, L. Hughes, B. Huntley, A. S. Van Jaarsveld, G. F. Midgley, L. Miles, M. A. Ortega-Huerta, A. T. Peterson, O. L. Phillips and S. E. Williams, *Nature*, 2004, **427**, 145–148.
- P. M. Cox, R. A. Betts, C. D. Jones, S. A. Spall and I. J. Totterdell, *Nature*, 2000, **408**, 184–187.
- Z. Y. Sun, T. Ma, H. C. Tao, Q. Fan and B. X. Han, *Chem*, 2017, **3**, 560–587.
- S. Garg, M. Li, A. Z. Weber, L. Ge, L. Li, V. Rudolph, G. Wang and T. E. Rufford, *J. Mater. Chem. A*, 2020, **8**, 1511–1544.
- S. Nitopi, E. Bertheussen, S. B. Scott, X. Liu, A. K. Engstfeld, S. Horch, B. Seger, I. E. L. Stephens, K. Chan, C. Hahn, J. K. Nørskov, T. F. Jaramillo and I. Chorkendorff, *Chem. Rev.*, 2019, **119**, 7610–7672.
- L. Fan, C. Xia, F. Yang, J. Wang, H. Wang and Y. Lu, *Sci. Adv.*, 2020, **6**, eaay3111.
- R. Kortlever, J. Shen, K. J. Schouten, F. Calle-Vallejo and M. T. Koper, *J. Phys. Chem. Lett.*, 2015, **6**, 4073–4082.
- D. F. Gao, R. M. Aran-Ais, H. S. Jeon and B. Roldán Cuenya, *Nat. Catal.*, 2019, **2**, 198–210.
- C. Du, X. Wang, W. Chen, S. Feng, J. Wen and Y. A. Wu, *Mater. Today Adv.*, 2020, **6**, 24.
- F. P. Pan and Y. Yang, *Energy Environ. Sci.*, 2020, **13**, 2275–2309.
- Y. Hori, K. Kikuchi and S. Suzuki, *Chem. Lett.*, 1985, 1695–1698, DOI: 10.1246/cl.1985.1695.
- R. Kortlever, I. Peters, C. Balemans, R. Kas, Y. Kwon, G. Mul and M. T. Koper, *Chem. Commun.*, 2016, **52**, 10229–10232.
- J. L. Qiao, Y. Y. Liu, F. Hong and J. J. Zhang, *Chem. Soc. Rev.*, 2014, **43**, 631–675.
- J. Hawecker, J.-M. Lehn and R. Ziessel, *J. Chem. Soc., Chem. Commun.*, 1984, 328–330, DOI: 10.1039/c39840000328.
- D. Sun, X. Xu, Y. Qin, S. P. Jiang and Z. Shao, *ChemSusChem*, 2020, **13**, 39–58.
- A. Salehi-Khojin, H. R. M. Jhong, B. A. Rosen, W. Zhu, S. C. Ma, P. J. A. Kenis and R. I. Masel, *J. Phys. Chem. C*, 2013, **117**, 1627–1632.
- R. Reske, H. Mistry, F. Behafarid, B. Roldán Cuenya and P. Strasser, *J. Am. Chem. Soc.*, 2014, **136**, 6978–6986.
- H. Mistry, R. Reske, Z. Zeng, Z. J. Zhao, J. Greeley, P. Strasser and B. R. Cuenya, *J. Am. Chem. Soc.*, 2014, **136**, 16473–16476.
- H. Mistry, R. Reske, P. Strasser and B. Roldán Cuenya, *Catal. Today*, 2017, **288**, 30–36.
- H. S. Jeon, I. Sinev, F. Scholten, N. J. Divins, I. Zegkinoglou, L. Pielsticker and B. R. Cuenya, *J. Am. Chem. Soc.*, 2018, **140**, 9383–9386.
- Y. Sun, S. Gao, F. Lei and Y. Xie, *Chem. Soc. Rev.*, 2015, **44**, 623–636.
- H. Tao, Y. Gao, N. Talreja, F. Guo, J. Texter, C. Yan and Z. Sun, *J. Mater. Chem. A*, 2017, **5**, 7257–7284.
- W. Bi, X. Li, R. You, M. Chen, R. Yuan, W. Huang, X. Wu, W. Chu, C. Wu and Y. Xie, *Adv. Mater.*, 2018, **30**, e1706617.
- Z. Lin, B. R. Carvalho, E. Kahn, R. Lv, R. Rao, H. Terrones, M. A. Pimenta and M. Terrones, *2D Mater.*, 2016, **3**, 022002.
- H. Chen, A. D. Handoko, J. Xiao, X. Feng, Y. Fan, T. Wang, D. Legut, Z. W. Seh and Q. Zhang, *ACS Appl. Mater. Interfaces*, 2019, **11**, 36571–36579.
- F. P. García De Arquer, O. S. Bushuyev, P. De Luna, C.-T. Dinh, A. Seifitokaldani, M. I. Saidaminov, C.-S. Tan, L. N. Quan, A. Proppe, M. G. Kibria, S. O. Kelley, D. Sinton and E. H. Sargent, *Adv. Mater.*, 2018, **30**, 1802858.
- H. Cheng, S. Liu, J. Zhang, T. Zhou, N. Zhang, X. S. Zheng, W. Chu, Z. Hu, C. Wu and Y. Xie, *Nano Lett.*, 2020, **20**, 6097–6103.
- A. D. Handoko, F. X. Wei, Jenndy, B. S. Yeo and Z. W. Seh, *Nat. Catal.*, 2018, **1**, 922–934.
- K. F. Kalz, R. Kraehnert, M. Dvoyashkin, R. Dittmeyer, R. Glaser, U. Krewer, K. Reuter and J. D. Grunwaldt, *ChemCatChem*, 2017, **9**, 17–29.
- R. Zhou, X. Fan, X. Ke, J. Xu, X. Zhao, L. Jia, B. Pan, N. Han, L. Li, X. Liu, J. Luo, H. Lin and Y. Li, *Nano Lett.*, 2021, **21**, 4092–4098.
- Y. Lan, G. Niu, F. Wang, D. Cui and Z. Hu, *ACS Appl. Mater. Interfaces*, 2020, **12**, 36128–36136.
- W. Huang, D. Zhou, H. Yang, X. Liu and J. Luo, *ACS Appl. Energy Mater.*, 2021, **4**, 7492–7496.
- Q. Zhao, C. Zhang, R. Hu, Z. Du, J. Gu, Y. Cui, X. Chen, W. Xu, Z. Cheng, S. Li, B. Li, Y. Liu, W. Chen, C. Liu, J. Shang, L. Song and S. Yang, *ACS Nano*, 2021, **15**, 4927–4936.
- C. Cao, D. D. Ma, J. Jia, Q. Xu, X. T. Wu and Q. L. Zhu, *Adv. Mater.*, 2021, **33**, 2008631.
- M. Cokoja, C. Bruckmeier, B. Rieger, W. A. Herrmann and F. E. Kuhn, *Angew. Chem., Int. Ed.*, 2011, **50**, 8510–8537.
- W. J. Zhang, Y. Hu, L. B. Ma, G. Y. Zhu, P. Y. Zhao, X. L. Xue, R. P. Chen, S. Y. Yang, J. Ma, J. Liu and Z. Jin, *Nano Energy*, 2018, **53**, 808–816.
- H. Zhang, Y. Ma, F. J. Quan, J. J. Huang, F. L. Jia and L. Z. Zhang, *Electrochem. Commun.*, 2014, **46**, 63–66.

- 38 P. Su, W. Xu, Y. Qiu, T. Zhang, X. Li and H. Zhang, *ChemSusChem*, 2018, **11**, 848–853.
- 39 L. C. Yi, J. X. Chen, P. Shao, J. H. Huang, X. X. Peng, J. W. Li, G. X. Wang, C. Zhang and Z. H. Wen, *Angew. Chem., Int. Ed.*, 2020, **59**, 20112–20119.
- 40 H. Yang, N. Han, J. Deng, J. Wu, Y. Wang, Y. Hu, P. Ding, Y. Li, Y. Li and J. Lu, *Adv. Energy Mater.*, 2018, **8**, 1801536.
- 41 C. J. Peng, G. Zeng, D. D. Ma, C. Cao, S. Zhou, X. T. Wu and Q. L. Zhu, *ACS Appl. Mater. Interfaces*, 2021, **13**, 20589–20597.
- 42 C.-Y. Lee, Y. Zhao, C. Wang, D. R. G. Mitchell and G. G. Wallace, *Sustainable Energy Fuels*, 2017, **1**, 1023–1027.
- 43 T. T. Zhang, X. F. Li, Y. L. Qiu, P. P. Su, W. B. Xu, H. X. Zhong and H. M. Zhang, *J. Catal.*, 2018, **357**, 154–162.
- 44 J. Pan, Y. M. Sun, P. L. Deng, F. Yang, S. H. Chen, Q. Zhou, H. S. Park, H. F. Liu and B. Y. Xia, *Appl. Catal., B*, 2019, **255**, 117736.
- 45 L. Dai, Q. Qin, P. Wang, X. Zhao, C. Hu, P. Liu, R. Qin, M. Chen, D. Ou, C. Xu, S. Mo, B. Wu, G. Fu, P. Zhang and N. Zheng, *Sci. Adv.*, 2017, **3**, e1701069.
- 46 W. Zhang, Q. Qin, L. Dai, R. Qin, X. Zhao, X. Chen, D. Ou, J. Chen, T. T. Chuong, B. Wu and N. Zheng, *Angew. Chem., Int. Ed.*, 2018, **57**, 9475–9479.
- 47 W. Deng, L. Zhang, L. Li, S. Chen, C. Hu, Z. J. Zhao, T. Wang and J. Gong, *J. Am. Chem. Soc.*, 2019, **141**, 2911–2915.
- 48 M. F. Baruch, J. E. Pander, J. L. White and A. B. Bocarsly, *ACS Catal.*, 2015, **5**, 3148–3156.
- 49 A. Dutta, A. Kuzume, V. Kaliginedi, M. Rahaman, I. Sinev, M. Ahmadi, B. Roldán Cuenya, S. Vesztergom and P. Broekmann, *Nano Energy*, 2018, **53**, 828–840.
- 50 S. Gao, Y. Lin, X. Jiao, Y. Sun, Q. Luo, W. Zhang, D. Li, J. Yang and Y. Xie, *Nature*, 2016, **529**, 68–71.
- 51 J. Jia, C. Qian, Y. Dong, Y. F. Li, H. Wang, M. Ghoussoub, K. T. Butler, A. Walsh and G. A. Ozin, *Chem. Soc. Rev.*, 2017, **46**, 4631–4644.
- 52 S. Gao, X. Jiao, Z. Sun, W. Zhang, Y. Sun, C. Wang, Q. Hu, X. Zu, F. Yang, S. Yang, L. Liang, J. Wu and Y. Xie, *Angew. Chem., Int. Ed.*, 2016, **55**, 698–702.
- 53 F. Li, L. Chen, G. P. Knowles, D. R. MacFarlane and J. Zhang, *Angew. Chem., Int. Ed.*, 2017, **56**, 505–509.
- 54 S. Gao, Z. Sun, W. Liu, X. Jiao, X. Zu, Q. Hu, Y. Sun, T. Yao, W. Zhang, S. Wei and Y. Xie, *Nat. Commun.*, 2017, **8**, 14503.
- 55 Z. Geng, X. Kong, W. Chen, H. Su, Y. Liu, F. Cai, G. Wang and J. Zeng, *Angew. Chem., Int. Ed.*, 2018, **57**, 6054–6059.
- 56 N. Sakai, Y. Ebina, K. Takada and T. Sasaki, *J. Am. Chem. Soc.*, 2004, **126**, 5851–5858.
- 57 X. Y. Chen, Y. Zhou, Q. Liu, Z. D. Li, J. G. Liu and Z. G. Zou, *ACS Appl. Mater. Interfaces*, 2012, **4**, 3372–3377.
- 58 Y. Zhou, Z. P. Tian, Z. Y. Zhao, Q. Liu, J. H. Kou, X. Y. Chen, J. Gao, S. C. Yan and Z. G. Zou, *ACS Appl. Mater. Interfaces*, 2011, **3**, 3594–3601.
- 59 Q. Liu, D. Wu, Y. Zhou, H. Su, R. Wang, C. Zhang, S. Yan, M. Xiao and Z. Zou, *ACS Appl. Mater. Interfaces*, 2014, **6**, 2356–2361.
- 60 S. Y. Zhu, S. J. Liang, J. H. Bi, M. H. Liu, L. M. Zhou, L. Wu and X. X. Wang, *Green Chem.*, 2016, **18**, 1355–1363.
- 61 S. Gao, B. Gu, X. Jiao, Y. Sun, X. Zu, F. Yang, W. Zhu, C. Wang, Z. Feng, B. Ye and Y. Xie, *J. Am. Chem. Soc.*, 2017, **139**, 3438–3445.
- 62 M. Asadi, B. Kumar, A. Behranginia, B. A. Rosen, A. Baskin, N. Repnin, D. Pisasale, P. Phillips, W. Zhu, R. Haasch, R. F. Klie, P. Kral, J. Abiade and A. Salehi-Khojin, *Nat. Commun.*, 2014, **5**, 4470.
- 63 M. Asadi, K. Kim, C. Liu, A. V. Addepalli, P. Abbasi, P. Yasaei, P. Phillips, A. Behranginia, J. M. Cerrato, R. Haasch, P. Zapol, B. Kumar, R. F. Klie, J. Abiade, L. A. Curtiss and A. Salehi-Khojin, *Science*, 2016, **353**, 467–470.
- 64 P. Abbasi, M. Asadi, C. Liu, S. Sharifi-Asl, B. Sayahpour, A. Behranginia, P. Zapol, R. Shahbazian-Yassar, L. A. Curtiss and A. Salehi-Khojin, *ACS Nano*, 2017, **11**, 453–460.
- 65 J. Xu, X. Li, W. Liu, Y. Sun, Z. Ju, T. Yao, C. Wang, H. Ju, J. Zhu, S. Wei and Y. Xie, *Angew. Chem., Int. Ed.*, 2017, **56**, 9121–9125.
- 66 X. Hong, K. R. Chan, C. Tsai and J. K. Nørskov, *ACS Catal.*, 2016, **6**, 4428–4437.
- 67 X. Jiao, Z. Chen, X. Li, Y. Sun, S. Gao, W. Yan, C. Wang, Q. Zhang, Y. Lin, Y. Luo and Y. Xie, *J. Am. Chem. Soc.*, 2017, **139**, 7586–7594.
- 68 X. D. Li, Y. F. Sun, J. Q. Xu, Y. J. Shao, J. Wu, X. L. Xu, Y. Pan, H. X. Ju, J. F. Zhu and Y. Xie, *Nat. Energy*, 2019, **4**, 690–699.
- 69 X. Jiao, X. Li, X. Jin, Y. Sun, J. Xu, L. Liang, H. Ju, J. Zhu, Y. Pan, W. Yan, Y. Lin and Y. Xie, *J. Am. Chem. Soc.*, 2017, **139**, 18044–18051.
- 70 L. D. Ramírez-Valencia, E. Bailón-García, F. Carrasco-Marín and A. F. Pérez-Cadenas, *Catalysts*, 2021, **11**, 351.
- 71 H. Jin, C. Guo, X. Liu, J. Liu, A. Vasileff, Y. Jiao, Y. Zheng and S. Z. Qiao, *Chem. Rev.*, 2018, **118**, 6337–6408.
- 72 G. L. Chai and Z. X. Guo, *Chem. Sci.*, 2016, **7**, 1268–1275.
- 73 H. Wang and J. M. Lee, *J. Mater. Chem. A*, 2020, **8**, 10604–10624.
- 74 Y. Gogotsi and B. Anasori, *ACS Nano*, 2019, **13**, 8491–8494.
- 75 J. L. Hart, K. Hantanasirisakul, A. C. Lang, B. Anasori, D. Pinto, Y. Pivak, J. T. Van Omme, S. J. May, Y. Gogotsi and M. L. Taheri, *Nat. Commun.*, 2019, **10**, 522.
- 76 N. Li, X. Chen, W. J. Ong, D. R. MacFarlane, X. Zhao, A. K. Cheetham and C. Sun, *ACS Nano*, 2017, **11**, 10825–10833.
- 77 A. D. Handoko, K. H. Khoo, T. L. Tan, H. M. Jin and Z. W. Seh, *J. Mater. Chem. A*, 2018, **6**, 21885–21890.
- 78 Y. Xiao and W. Zhang, *Nanoscale*, 2020, **12**, 7660–7673.
- 79 H. Furukawa, K. E. Cordova, M. O’Keeffe and O. M. Yaghi, *Science*, 2013, **341**, 1230444.
- 80 X. X. Tang, L. Zhao, W. W. Sun and Y. Wang, *J. Power Sources*, 2020, **477**, 228919.

- 81 M. K. Lee, M. Shokouhimehr, S. Y. Kim and H. W. Jang, *Adv. Energy Mater.*, 2021, 2003990, DOI: 10.1002/aenm.202003990.
- 82 Q. Y. Cui, G. Q. Qin, W. H. Wang, K. R. Geethalakshmi, A. J. Du and Q. Sun, *Appl. Surf. Sci.*, 2020, **500**, 143993.
- 83 H. Zhong, M. Ghorbani-Asl, K. H. Ly, J. Zhang, J. Ge, M. Wang, Z. Liao, D. Makarov, E. Zschech, E. Brunner, I. M. Weidinger, J. Zhang, A. V. Krasheninnikov, S. Kaskel, R. Dong and X. Feng, *Nat. Commun.*, 2020, **11**, 1409.
- 84 S. Lin, C. S. Diercks, Y. B. Zhang, N. Kornienko, E. M. Nichols, Y. Zhao, A. R. Paris, D. Kim, P. Yang, O. M. Yaghi and C. J. Chang, *Science*, 2015, **349**, 1208–1213.
- 85 Y. Zhang, L. Li, S. X. Guo, X. Zhang, F. Li, A. M. Bond and J. Zhang, *ChemSusChem*, 2020, **13**, 59–77.
- 86 D. Voiry, H. S. Shin, K. P. Loh and M. Chhowalla, *Nat. Rev. Chem.*, 2018, **2**, 0105.
- 87 S. Gao, Z. Sun, W. Liu, X. Jiao, X. Zu, Q. Hu, Y. Sun, T. Yao, W. Zhang, S. Wei and Y. Xie, *Nat. Commun.*, 2017, **8**, 14503.
- 88 S. Chen, H. Wang, Z. Kang, S. Jin, X. Zhang, X. Zheng, Z. Qi, J. Zhu, B. Pan and Y. Xie, *Nat. Commun.*, 2019, **10**, 788.
- 89 C. Peng, G. Luo, J. Zhang, M. Chen, Z. Wang, T. K. Sham, L. Zhang, Y. Li and G. Zheng, *Nat. Commun.*, 2021, **12**, 1580.
- 90 S. Siahrostami, K. Jiang, M. Karamad, K. R. Chan, H. T. Wang and J. Norskov, *ACS Sustainable Chem. Eng.*, 2017, **5**, 11080–11085.
- 91 H. Terrones, R. Lv, M. Terrones and M. S. Dresselhaus, *Rep. Prog. Phys.*, 2012, **75**, 062501.
- 92 Z. Y. Shu, G. Y. Ye, J. Wang, S. Q. Liu, Z. He, W. W. Zhu, B. Liu and M. Liu, *Ionics*, 2021, **27**, 3089–3098.
- 93 S. Back, M. S. Yeom and Y. Jung, *ACS Catal.*, 2015, **5**, 5089–5096.
- 94 F. Li, M. Xue, J. Li, X. Ma, L. Chen, X. Zhang, D. R. MacFarlane and J. Zhang, *Angew. Chem., Int. Ed.*, 2017, **56**, 14718–14722.
- 95 S. Liu, H. Tao, L. Zeng, Q. Liu, Z. Xu, Q. Liu and J. L. Luo, *J. Am. Chem. Soc.*, 2017, **139**, 2160–2163.
- 96 J. Wu, M. Liu, P. P. Sharma, R. M. Yadav, L. Ma, Y. Yang, X. Zou, X. D. Zhou, R. Vajtai, B. I. Yakobson, J. Lou and P. M. Ajayan, *Nano Lett.*, 2016, **16**, 466–470.
- 97 Y. Liu, J. Zhao and Q. Cai, *Phys. Chem. Chem. Phys.*, 2016, **18**, 5491–5498.
- 98 N. Srekanth, M. A. Nazrulla, T. V. Vineesh, K. Sailaja and K. L. Phani, *Chem. Commun.*, 2015, **51**, 16061–16064.
- 99 C. M. Zhao, Y. Wang, Z. J. Li, W. X. Chen, Q. Xu, D. S. He, D. S. Xi, Q. H. Zhang, T. W. Yuan, Y. T. Qu, J. Yang, F. Y. Zhou, Z. K. Yang, X. Q. Wang, J. Wang, J. Luo, Y. F. Li, H. H. Duan, Y. Wu and Y. D. Li, *Joule*, 2019, **3**, 584–594.
- 100 C. Zhang, S. Yang, J. Wu, M. Liu, S. Yazdi, M. Ren, J. Sha, J. Zhong, K. Nie, A. S. Jalilov, Z. Li, H. Li, B. I. Yakobson, Q. Wu, E. Ringe, H. Xu, P. M. Ajayan and J. M. Tour, *Adv. Energy Mater.*, 2018, **8**, 1703487.
- 101 J. Wang, X. Huang, S. Xi, J. M. Lee, C. Wang, Y. Du and X. Wang, *Angew. Chem., Int. Ed.*, 2019, **58**, 13532–13539.
- 102 Q. Li, W. L. Zhu, J. J. Fu, H. Y. Zhang, G. Wu and S. H. Sun, *Nano Energy*, 2016, **24**, 1–9.
- 103 K. Jiang, S. Siahrostami, T. T. Zheng, Y. F. Hu, S. Hwang, E. Stavitski, Y. D. Peng, J. Dynes, M. Gangisetty, D. Su, K. Attenkofer and H. T. Wang, *Energy Environ. Sci.*, 2018, **11**, 893–903.
- 104 F. Li, G. H. Gu, C. Choi, P. Kolla, S. Hong, T. S. Wu, Y. L. Soo, J. Masa, S. Mukerjee, Y. S. Jung, J. S. Qiu and Z. Y. Sun, *Appl. Catal., B*, 2020, **277**, 119241.
- 105 M. G. Kibria, J. P. Edwards, C. M. Gabardo, C. T. Dinh, A. Seifitokaldani, D. Sinton and E. H. Sargent, *Adv. Mater.*, 2019, **31**, e1807166.
- 106 K. Liu, W. A. Smith and T. Burdyny, *ACS Energy Lett.*, 2019, **4**, 639–643.
- 107 T. Haas, R. Krause, R. Weber, M. Demler and G. Schmid, *Nat. Catal.*, 2018, **1**, 32–39.
- 108 H. Zhong, M. Ghorbani-Asl, K. H. Ly, J. Zhang, J. Ge, M. Wang, Z. Liao, D. Makarov, E. Zschech, E. Brunner, I. M. Weidinger, J. Zhang, A. V. Krasheninnikov, S. Kaskel, R. Dong and X. Feng, *Nat. Commun.*, 2020, **11**, 1409.
- 109 J. J. Velasco-Velez, R. V. Mom, L. E. Sandoval-Diaz, L. J. Falling, C. H. Chuang, D. Gao, T. E. Jones, Q. Zhu, R. Arrigo, B. Roldán Cuenya, A. Knop-Gericke, T. Lunkenbein and R. Schlogl, *ACS Energy Lett.*, 2020, **5**, 2106–2111.
- 110 J. B. Vennekoetter, R. Sengpiel and M. Wessling, *Chem. Eng. J.*, 2019, **364**, 89–101.
- 111 F. P. Garcia de Arquer, C. T. Dinh, A. Ozden, J. Wicks, C. McCallum, A. R. Kirmani, D. H. Nam, C. Gabardo, A. Seifitokaldani, X. Wang, Y. C. Li, F. Li, J. Edwards, L. J. Richter, S. J. Thorpe, D. Sinton and E. H. Sargent, *Science*, 2020, **367**, 661–666.
- 112 L. Li, A. Ozden, S. Guo, F. P. García De Arquer, C. Wang, M. Zhang, J. Zhang, H. Jiang, W. Wang, H. Dong, D. Sinton, E. H. Sargent and M. Zhong, *Nat. Commun.*, 2021, **12**, 5223.
- 113 D. Kim, W. Choi, H. W. Lee, S. Y. Lee, Y. Choi, D. K. Lee, W. Kim, J. Na, U. Lee, Y. J. Hwang and D. H. Won, *ACS Energy Lett.*, 2021, **6**, 3488–3495.
- 114 C. Xia, P. Zhu, Q. Jiang, Y. Pan, W. T. Liang, E. Stavitski, H. N. Alshareef and H. T. Wang, *Nat. Energy*, 2019, **4**, 776–785.
- 115 P. Zhu, C. Xia, C.-Y. Liu, K. Jiang, G. Gao, X. Zhang, Y. Xia, Y. Lei, H. N. Alshareef, T. P. Senftle and H. Wang, *Proc. Natl. Acad. Sci. U. S. A.*, 2021, **118**, e2010868118.
- 116 L. Fan, C. Xia, P. Zhu, Y. Lu and H. Wang, *Nat. Commun.*, 2020, **11**, DOI: 10.1038/s41467-020-17403-1.
- 117 S. F. Tan, S. W. Chee, Z. Baraissov, H. M. Jin, T. L. Tan and U. Mirsaidov, *Adv. Funct. Mater.*, 2019, **29**, 1903242.
- 118 J. Yin, S. Y. Shan, L. F. Yang, D. Mott, O. Malis, V. Petkov, F. Cai, M. S. Ng, J. Luo, B. H. Chen, M. Engelhard and C. J. Zhong, *Chem. Mater.*, 2012, **24**, 4662–4674.
- 119 T. Avanesian, S. Dai, M. J. Kale, G. W. Graham, X. Pan and P. Christopher, *J. Am. Chem. Soc.*, 2017, **139**, 4551–4558.

- 120 W. Zhan, J. Wang, H. Wang, J. Zhang, X. Liu, P. Zhang, M. Chi, Y. Guo, Y. Guo, G. Lu, S. Sun, S. Dai and H. Zhu, *J. Am. Chem. Soc.*, 2017, **139**, 8846–8854.
- 121 M. D. Argyle and C. H. Bartholomew, *Catalysts*, 2015, **5**, 145–269.
- 122 S. Popovic, M. Smiljanic, P. Jovanovic, J. Vavra, R. Buonsanti and N. Hodnik, *Angew. Chem., Int. Ed.*, 2020, **59**, 14736–14746.
- 123 D. Karapinar, N. T. Huan, N. Ranjbar Sahraie, J. Li, D. Wakerley, N. Touati, S. Zanna, D. Taverna, L. H. Galvão Tizei, A. Zitolo, F. Jaouen, V. Mougél and M. Fontecave, *Angew. Chem., Int. Ed.*, 2019, **58**, 15098–15103.
- 124 S. W. Chee, T. Lunkenbein, R. Schlogl and B. R. Cuenya, *J. Phys.: Condens. Matter*, 2021, **33**, 153001.
- 125 Y. Zhu, J. Wang, H. Chu, Y.-C. Chu and H. M. Chen, *ACS Energy Lett.*, 2020, **5**, 1281–1291.
- 126 M. A. Bañares, *Catal. Today*, 2005, **100**, 71–77.
- 127 A. Preet and T. E. Lin, *Catalysts*, 2021, **11**, 594.
- 128 F. D. Mayer, P. Hosseini-Benhangi, C. M. Sanchez-Sanchez, E. Asselin and E. L. Gyenge, *Commun. Chem.*, 2020, **3**, 155.
- 129 S. Narayanaru, J. Chinnaiyah, K. L. Phani and F. Scholz, *Electrochim. Acta*, 2018, **264**, 269–274.
- 130 C. I. Shaughnessy, D. T. Jantz and K. C. Leonard, *J. Mater. Chem. A*, 2017, **5**, 22743–22749.
- 131 T. Kai, M. Zhou, Z. Duan, G. A. Henkelman and A. J. Bard, *J. Am. Chem. Soc.*, 2017, **139**, 18552–18557.
- 132 D. T. Jantz and K. C. Leonard, *Ind. Eng. Chem. Res.*, 2018, **57**, 7431–7440.
- 133 Y. Kim, A. Jo, Y. Ha, Y. Lee, D. Lee, Y. Lee and C. Lee, *Electroanalysis*, 2018, **30**, 2861–2868.
- 134 N. Srekanth and K. L. Phani, *Chem. Commun.*, 2014, **50**, 11143–11146.
- 135 L. I. Stephens, N. A. Payne and J. Mauzeroll, *Anal. Chem.*, 2020, **92**, 3958–3963.
- 136 D. Polcari, P. Dauphin-Ducharme and J. Mauzeroll, *Chem. Rev.*, 2016, **116**, 13234–13278.
- 137 D. Wang and L. J. Wan, *J. Phys. Chem. C*, 2007, **111**, 16109–16130.
- 138 H. Feng, X. Xu, Y. Du and S. X. Dou, *Electrochem. Energy Rev.*, 2021, **4**, 249–268.
- 139 Y. Liang, J. H. K. Pfisterer, D. McLaughlin, C. Csoklich, L. Seidl, A. S. Bandarenka and O. Schneider, *Small Methods*, 2019, **3**, 1800387.
- 140 C. F. Tsang, A. C. Javier, Y.-G. Kim, J. H. Baricuatro, K. D. Cummins, J. Kim, G. Jerkiewicz, J. C. Hemminger and M. P. Soriaga, *J. Electrochem. Soc.*, 2018, **165**, J3350–J3354.
- 141 Y.-G. Kim and M. P. Soriaga, *J. Electroanal. Chem.*, 2014, **734**, 7–9.
- 142 Y.-G. Kim, J. H. Baricuatro, A. Javier, J. M. Gregoire and M. P. Soriaga, *Langmuir*, 2014, **30**, 15053–15056.
- 143 G. H. Simon, C. S. Kley and B. Roldán Cuenya, *Angew. Chem., Int. Ed.*, 2021, **60**, 2561–2568.
- 144 M. R. Nellist, F. A. L. Laskowski, J. J. Qiu, H. Hajibabaei, K. Sivula, T. W. Hamann and S. W. Boettcher, *Nat. Energy*, 2018, **3**, 46–52.
- 145 P. Grosse, D. Gao, F. Scholten, I. Sinev, H. Mistry and B. Roldán Cuenya, *Angew. Chem., Int. Ed.*, 2018, **57**, 6192–6197.
- 146 Y. A. Wu, Z. Yin, M. Farmand, Y.-S. Yu, D. A. Shapiro, H.-G. Liao, W.-I. Liang, Y.-H. Chu and H. Zheng, *Sci. Rep.*, 2017, **7**, 42527.
- 147 Y. A. Wu, L. Li, Z. Li, A. Kinaci, M. K. Y. Chan, Y. Sun, J. R. Guest, I. McNulty, T. Rajh and Y. Liu, *ACS Nano*, 2016, **10**, 3738–3746.
- 148 Y. A. Wu, I. McNulty, C. Liu, K. C. Lau, Q. Liu, A. P. Paulikas, C. J. Sun, Z. H. Cai, J. R. Guest, Y. Ren, V. Stamenkovic, L. A. Curtiss, Y. Z. Liu and T. Rajh, *Nat. Energy*, 2019, **4**, 957–968.
- 149 B. D. A. Levin, D. Haiber, Q. Liu and P. A. Crozier, *Microsc. Microanal.*, 2020, **26**, 134–138.
- 150 K. Jungjohann and C. B. Carter, in *Transmission Electron Microscopy*, Springer International Publishing, 2016, pp. 17–80, DOI: 10.1007/978-3-319-26651-0_2.
- 151 M. J. Williamson, R. M. Tromp, P. M. Vereecken, R. Hull and F. M. Ross, *Nat. Mater.*, 2003, **2**, 532–536.
- 152 E. Fahrenkrug, D. H. Alsem, N. Salmon and S. Maldonado, *J. Electrochem. Soc.*, 2017, **164**, H358–H364.
- 153 X. Wang, K. Klingan, M. Klingenhof, T. Möller, J. Ferreira de Araújo, I. Martens, A. Bagger, S. Jiang, J. Rossmeisl, H. Dau and P. Strasser, *Nat. Commun.*, 2021, **12**, 794.
- 154 J. Vavra, T. H. Shen, D. Stoian, V. Tileli and R. Buonsanti, *Angew. Chem., Int. Ed.*, 2021, **60**, 1347–1354.
- 155 D. A. Shapiro, Y.-S. Yu, T. Tylliszczak, J. Cabana, R. Celestre, W. Chao, K. Kaznatcheev, A. L. D. Kilcoyne, F. Maia, S. Marchesini, Y. S. Meng, T. Warwick, L. L. Yang and H. A. Padmore, *Nat. Photonics*, 2014, **8**, 765–769.
- 156 J. Deng, D. J. Vine, S. Chen, Y. S. G. Nashed, Q. Jin, N. W. Phillips, T. Peterka, R. Ross, S. Vogt and C. J. Jacobsen, *Proc. Natl. Acad. Sci. U. S. A.*, 2015, **112**, 2314–2319.
- 157 T. H. Yoon, *Appl. Spectrosc. Rev.*, 2009, **44**, 91–122.
- 158 B. Rösner, S. Finizio, F. Koch, F. Döring, V. A. Guzenko, M. Langer, E. Kirk, B. Watts, M. Meyer, J. Loroña Ornelas, A. Späth, S. Stanescu, S. Swaraj, R. Belkhou, T. Ishikawa, T. F. Keller, B. Gross, M. Poggio, R. H. Fink, J. Raabe, A. Kleibert and C. David, *Optica*, 2020, **7**, 1602–1608.
- 159 S. Finizio, S. Mayr and J. Raabe, *J. Synchrotron Radiat.*, 2020, **27**, 1320–1325.
- 160 E. De Smit, I. Swart, J. F. Creemer, G. H. Hovelings, M. K. Gilles, T. Tylliszczak, P. J. Kooyman, H. W. Zandbergen, C. Morin, B. M. Weckhuysen and F. M. F. De Groot, *Nature*, 2008, **456**, 222–225.
- 161 D. Guay, J. Stewart-Ornstein, X. Zhang and A. P. Hitchcock, *Anal. Chem.*, 2005, **77**, 3479–3487.
- 162 A. Ulvestad, A. Singer, J. N. Clark, H. M. Cho, J. W. Kim, R. Harder, J. Maser, Y. S. Meng and O. G. Shpyrko, *Science*, 2015, **348**, 1344–1347.
- 163 I. Robinson and R. Harder, *Nat. Mater.*, 2009, **8**, 291–298.

- 164 M. Abuin, Y. Y. Kim, H. Runge, S. Kulkarni, S. Maier, D. Dzhigaev, S. Lazarev, L. Gelisio, C. Seitz, M. I. Richard, T. Zhou, V. Vonk, T. F. Keller, I. A. Vartanyants and A. Stierle, *ACS Appl. Nano Mater.*, 2019, **2**, 4818–4824.
- 165 R. Harder, M. Liang, Y. Sun, Y. Xia and I. K. Robinson, *New J. Phys.*, 2010, **12**, 035019.
- 166 W. Cha, A. Ulvestad, M. Allain, V. Chamard, R. Harder, S. J. Leake, J. Maser, P. H. Fuoss and S. O. Hruszkewycz, *Phys. Rev. Lett.*, 2016, **117**, 225501.
- 167 I. Robinson, J. Clark and R. Harder, *J. Opt.*, 2016, **18**, 054007.
- 168 A. Ulvestad, A. Tripathi, S. O. Hruszkewycz, W. Cha, S. M. Wild, G. B. Stephenson and P. H. Fuoss, *Phys. Rev. B*, 2016, **93**, 184105.
- 169 R. P. Janssonius, L. M. Reid, C. N. Virca and C. P. Berlinguette, *ACS Energy Lett.*, 2019, **4**, 980–986.
- 170 S. Fernández, L. Gao, J. P. Hofmann, J. Carnis, S. Labat, G. A. Chahine, A. J. F. Van Hoof, M. W. G. M. (Tiny) Verhoeven, T. U. Schüllli, E. J. M. Hensen, O. Thomas and M.-I. Richard, *Nanoscale*, 2019, **11**, 331–338.
- 171 X. Zheng, B. Zhang, P. De Luna, Y. Liang, R. Comin, O. Voznyy, L. Han, F. P. García de Arquer, M. Liu, C. T. Dinh, T. Regier, J. J. Dynes, S. He, H. L. Xin, H. Peng, D. Prendergast, X. Du and E. H. Sargent, *Nat. Chem.*, 2018, **10**, 149–154.
- 172 J. Timoshenko and B. Roldán Cuenya, *Chem. Rev.*, 2021, **121**, 882–961.
- 173 M. Benfatto and C. Meneghini, in *Synchrotron Radiation*, Springer Berlin Heidelberg, 2015, ch. 7, pp. 213–240, DOI: 10.1007/978-3-642-55315-8_7.
- 174 J. J. Velasco-Velez, C. H. Wu, T. A. Pascal, L. F. Wan, J. Guo, D. Prendergast and M. Salmeron, *Science*, 2014, **346**, 831–834.
- 175 P. De Luna, R. Quintero-Bermudez, C.-T. Dinh, M. B. Ross, O. S. Bushuyev, P. Todorović, T. Regier, S. O. Kelley, P. Yang and E. H. Sargent, *Nat. Catal.*, 2018, **1**, 103–110.
- 176 J. Wang, C. Cheng, B. Huang, J. Cao, L. Li, Q. Shao, L. Zhang and X. Huang, *Nano Lett.*, 2021, **21**, 980–987.
- 177 L. P. Zhong, D. K. Chen and S. Zafeiratos, *Catal. Sci. Technol.*, 2019, **9**, 3851–3867.
- 178 K. Roy, L. Artiglia and J. A. van Bokhoven, *ChemCatChem*, 2018, **10**, 666–682.
- 179 A. Kolmakov, D. A. Dikin, L. J. Cote, J. Huang, M. K. Abyaneh, M. Amati, L. Gregoratti, S. Gunther and M. Kiskinova, *Nat. Nanotechnol.*, 2011, **6**, 651–657.
- 180 M. Favaro, H. Xiao, T. Cheng, W. A. Goddard 3rd, J. Yano and E. J. Crumlin, *Proc. Natl. Acad. Sci. U. S. A.*, 2017, **114**, 6706–6711.
- 181 J. Kim, H. Ha, W. H. Doh, K. Ueda, K. Mase, H. Kondoh, B. S. Mun, H. Y. Kim and J. Y. Park, *Nat. Commun.*, 2020, **11**, 5649.
- 182 R. Blume, D. Rosenthal, J. P. Tessonnier, H. N. Li, A. Knop-Gericke and R. Schlogl, *ChemCatChem*, 2015, **7**, 2871–2881.
- 183 U. Bergmann, J. Kern, R. W. Schoenlein, P. Wernet, V. K. Yachandra and J. Yano, *Nat. Rev. Phys.*, 2021, **3**, 264–282.
- 184 I. Matsuda and Y. Kubota, *Chem. Lett.*, 2021, **50**, 1336–1344.
- 185 G. Geloni, Z. Huang and C. Pellegrini, in *X-Ray Free Electron Lasers*, The Royal Society of Chemistry, 2017, pp. 1–44, DOI: 10.1039/9781782624097-00001.
- 186 C. Pellegrini, A. Marinelli and S. Reiche, *Rev. Mod. Phys.*, 2016, **88**, 015006.
- 187 J. Yang, R. Dettori, J. P. F. Nunes, N. H. List, E. Biasin, M. Centurion, Z. Chen, A. A. Cordones, D. P. Deponte, T. F. Heinz, M. E. Kozina, K. Ledbetter, M.-F. Lin, A. M. Lindenberg, M. Mo, A. Nilsson, X. Shen, T. J. A. Wolf, D. Donadio, K. J. Gaffney, T. J. Martinez and X. Wang, *Nature*, 2021, **596**, 531–535.
- 188 A. H. Zewail, *Science*, 1988, **242**, 1645–1653.
- 189 L. B. Hoch, P. Szymanski, K. K. Ghuman, L. He, K. Liao, Q. Qiao, L. M. Reyes, Y. Zhu, M. A. El-Sayed, C. V. Singh and G. A. Ozin, *Proc. Natl. Acad. Sci. U. S. A.*, 2016, **113**, E8011–E8020.
- 190 J. Sá and J. Szlachetko, *Catal. Lett.*, 2013, **144**, 197–203.
- 191 E. L. Clark, M. R. Singh, Y. Kwon and A. T. Bell, *Anal. Chem.*, 2015, **87**, 8013–8020.
- 192 E. L. Clark and A. T. Bell, *J. Am. Chem. Soc.*, 2018, **140**, 7012–7020.
- 193 T. Herl and F. M. Matysik, *ChemElectroChem*, 2020, **7**, 2498–2512.
- 194 M. Plodinec, H. C. Nerl, R. Farra, M. G. Willinger, E. Stotz, R. Schlogl and T. Lunkenbein, *Microsc. Microanal.*, 2020, **26**, 220–228.
- 195 J. Lu, X. Hua and Y. T. Long, *Analyst*, 2017, **142**, 691–699.
- 196 R. Kas, O. Ayemoba, N. J. Firet, J. Middelkoop, W. A. Smith and A. Cuesta, *ChemPhysChem*, 2019, **20**, 2904–2925.
- 197 F. Zaera, *Chem. Soc. Rev.*, 2014, **43**, 7624–7663.
- 198 A. J. Wain and M. A. O'Connell, *Adv. Phys.: X*, 2017, **2**, 188–209.
- 199 J. Heyes, M. Dunwell and B. Xu, *J. Phys. Chem. C*, 2016, **120**, 17334–17341.
- 200 C. Hess, *Chem. Soc. Rev.*, 2021, **50**, 3519–3564.
- 201 Y. Wang, Y. Li, J. Liu, C. Dong, C. Xiao, L. Cheng, H. Jiang, H. Jiang and C. Li, *Angew. Chem.*, 2021, **60**, 7681–7685.
- 202 J. F. Li, Y. F. Huang, Y. Ding, Z. L. Yang, S. B. Li, X. S. Zhou, F. R. Fan, W. Zhang, Z. Y. Zhou, D. Y. Wu, B. Ren, Z. L. Wang and Z. Q. Tian, *Nature*, 2010, **464**, 392–395.
- 203 Z. Zhang, S. Sheng, R. Wang and M. Sun, *Anal. Chem.*, 2016, **88**, 9328–9346.
- 204 N. Kumar, C. S. Wondergem, A. J. Wain and B. M. Weckhuysen, *J. Phys. Chem. Lett.*, 2019, **10**, 1669–1675.
- 205 E. Pérez-Gallent, G. Marcandalli, M. C. Figueiredo, F. Calle-Vallejo and M. T. M. Koper, *J. Am. Chem. Soc.*, 2017, **139**, 16412–16419.

- 206 S. Lee, D. Kim and J. Lee, *Angew. Chem.*, 2015, **127**, 14914–14918.
- 207 A. S. Varela, W. Ju, T. Reier and P. Strasser, *ACS Catal.*, 2016, **6**, 2136–2144.
- 208 D. Bohra, J. H. Chaudhry, T. Burdyny, E. A. Pidko and W. A. Smith, *Energy Environ. Sci.*, 2019, **12**, 3380–3389.
- 209 Y. Wang, M. Hatakeyama, K. Ogata, M. Wakabayashi, F. Jin and S. Nakamura, *Phys. Chem. Chem. Phys.*, 2015, **17**, 23521–23531.
- 210 S. Banerjee, X. Han and V. S. Thoi, *ACS Catal.*, 2019, **9**, 5631–5637.
- 211 L. Zhang, Z. C. Wei, S. Thanneeru, M. Meng, M. Kruzyk, G. Ung, B. Liu and J. He, *Angew. Chem., Int. Ed.*, 2019, **58**, 15834–15840.
- 212 A. K. Buckley, M. Lee, T. Cheng, R. V. Kazantsev, D. M. Larson, W. A. Goddard III, F. D. Toste and F. M. Toma, *J. Am. Chem. Soc.*, 2019, **141**, 7355–7364.
- 213 A. Wagner, C. D. Sahm and E. Reisner, *Nat. Catal.*, 2020, **3**, 775–786.
- 214 A. Khorshidi, J. Violet, J. Hashemi and A. A. Peterson, *Nat. Catal.*, 2018, **1**, 263–268.
- 215 Z. Dai, L. Liu and Z. Zhang, *Adv. Mater.*, 2019, **31**, 1805417.