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Topology and porosity control of metal–organic frameworks through linker functionalization†

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Tetrapotic organic linkers have been extensively used in Zr-based metal–organic frameworks (MOFs) where diverse topologies have been observed. Achieving meticulous control over the topologies to tune the pore sizes and shapes of the resulting materials, however, remains a great challenge. Herein, by introducing substituents to the backbone of tetrapotic linkers to affect the linker conformation, phase-pure Zr-MOFs with different topologies and porosity were successfully obtained under the same synthetic conditions. The conversion of CO₂ to valuable cyclic carbonates is a promising route for the mitigation of the greenhouse gas. Owing to the presence of substrate accessible Lewis acidic Zr(IV) sites in the 8-connected Zr₆ nodes, the Zr-MOFs in this study have been investigated as heterogeneous acid catalysts for CO₂ cycloaddition to styrene oxide. The MOFs exhibited drastically different catalytic activities depending on their distinct pore structures. Compared to previously reported MOF materials, a superior catalytic activity was observed with the mesoporous NU-1008, giving an almost 100% conversion under mild conditions.

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Introduction

Metal–organic frameworks (MOFs),^{1–3} a class of crystalline porous materials constructed from metal nodes and organic ligands by coordination bonds, have found a wide range of applications such as gas storage and separation,^{4–11} catalysis,^{12–16} magnetism,^{17,18} drug delivery,^{19–23} small-molecule recognition²⁴ and chemical sensing.^{25,26} This functional diversity is largely due to their versatile structural tunability where a variety of components can be combined to generate an almost indefinite number of novel materials. However, elucidating the parameters that dictate the MOF structures, including their topology and porosity, is non-trivial. Considering Zr₆ cluster-based MOFs as an example, although linkers of lower connectivity generally favor one particular topology,^{27–29} more topologies can be accessed from linkers with higher connectivity^{30,31}

due to their tendency to undergo conformational changes, resulting in challenges in predicting MOF structures.^{32–34}

Since the first reports of Zr-MOFs with tetrapotic linkers,^{35–37} several topologies have been observed, including **ftw**,^{36,38–41} **csq**,^{35–37,42} **she**,⁴³ **shp**,⁴⁴ **scu**,^{32,45} **flu**,^{27,46,47} **ith**,²⁷ **sqc**,⁴⁸ **lvt**⁴⁹ and **stp**⁵⁰ which can be attributed to the linker conformation adjustments under different synthetic conditions. This conformational change can be induced by controlling one of the many factors, such as the modulating reagent, concentration, metal salt, temperature and solvent. For example, a **csq**-net MOF (NU-1000)^{42,51} is produced when using benzoic acid as the modulator whereas a **scu**-net MOF (NU-901)³² is obtained using 4-amino-benzoic acid with the same 1,3,6,8-tetrakis(*p*-benzoic acid)pyrene (TBAPy) linker and Zr precursor. Using a tetrapotic linker with arms of a high degree of rotational freedom, such as tetracarboxyphenylporphyrin (TCPP), multiple topologies have been observed.^{1,35–37,43,44,48} A similar phenomenon was also observed when we explored the isorecticular tetracarboxylate Zr-MOF **csq**-net NU-100X series for enzyme immobilization where MOFs with **ftw** topology were produced instead when using linkers with longer arms. Therefore, our group concluded that the torsion angle between the planar “backbone” (benzene, pyrene, porphyrin) and the arms (carboxylic acid) had a critical influence on the framework topology. Specifically, if the torsion angle was close to 60°, the **csq** topology was favored, whereas the **ftw** topology formed if the torsion angle was close to 0°.⁵²

In addition to using organic linkers with higher rotational freedom, installing functional groups to the organic linkers

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affords another effective strategy to affect the linker conformation, thereby the material topology. For instance, Yaghi and co-workers successfully synthesized **qom**-net MOF-177 series with uncommon topologies (**pyr** and **rtl**) by introducing functional groups.⁵³ Recently, Zhou and co-workers introduced steric hindrance into the biphenyl-3,3',5,5'-tetra(phenyl-4-carboxylic acid) (TCPB) linker to affect linker conformation, which subsequently effect the topology of the resulting Zr-tetracarboxylate MOFs. However, accompanying the original structure were mixed phase materials that formed as byproducts, most likely due to the introduction of the bulky substituents.⁵⁴ In high agreement with experimental results, computational studies compared the energy of MOFs with different substituents to demonstrate the influence of introduced steric hindrance on linker conformation and MOF topology.⁵⁵ With these design rules and challenges in mind, we set out to tune the steric hindrance of the organic linkers in a stepwise fashion, aiming to achieve topology control in the resulting series of Zr-MOFs.

Results and discussion

Introduction of steric hindrance

In addition to the 1,2,4,5-tetrakis(4-carboxyphenyl)benzene (TCPB) linker (Fig. 1, L1), two more linkers were designed. A nitro group (Fig. 1, L2) and two bromo groups (Fig. 1, L3) were installed on the central benzene respectively to introduce steric hindrance to affect the rotational freedom of the peripheral benzoate arms and in turn to control the topology of the resulting MOFs. To ensure consistency and to eliminate variability from factors such as solvent, temperature and modulator on the topology of the resulting MOFs, the syntheses with the three TCPB-based linkers were carried out under identical solvothermal conditions (ESI†).

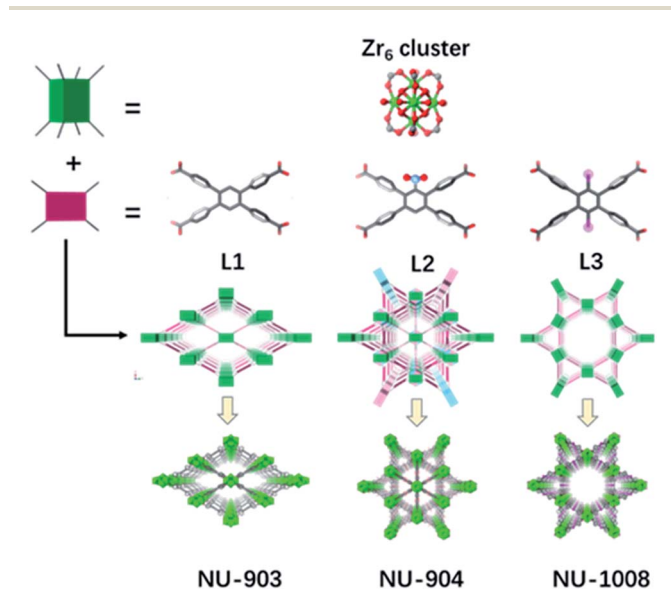


Fig. 1 Schematic representation of the construction of NU-903, NU-904 and NU-1008. Atom colour scheme: C, grey; N, blue; O, red; Br, pink; Zr, green polyhedra. H atoms are omitted for clarity.

Topology and porosity diversity of the as-synthesized MOFs

As previously reported by Stock and co-workers,⁴⁵ Zr₆ nodes and the L1 linker form NU-903 with an **scu** topology with characteristic diamond-shaped channels along the *c* axis (Fig. 1). NU-903 is isorecticular to TBAPy-based NU-901³² and TCPB-based NU-902.⁵⁶

The crystals of the resulting MOF with linker L2 are oval in shape (Fig. 3b). Single crystal X-ray diffraction analysis (Table S1, ESI†) revealed that NU-904 crystallized in the *P2/m* space group ($a = 19.64(1) \text{ \AA}$, $b = 12.63(1) \text{ \AA}$, $c = 19.635(4) \text{ \AA}$ and $\beta = 119.994^\circ$ (2) at 200 K) with the formula as $\text{Zr}_6(\mu\text{-O})_4(\mu\text{-OH})_4(\text{HCOO})_{1.5}(\text{H}_2\text{O})_{2.5}(\text{OH})_{2.5}(\text{L2})_2$. The 3D structure consists of 8-connected Zr₆ nodes and mononitro-substituted TCPB linkers, yielding a rare **scu** topology.³² The single crystal structure of NU-904 is characteristic of a reticular-merohedral twin structure in which three orientations stack together along the *b* axis, twisting 60° from each other (Fig. 2). The reticular twins of three **scu**-net components gave rise to the overall 6-fold symmetry in the structure. Topologically, the average structure of the threefold twinned NU-904 is based on the highly connected 4,12-**shp** net.^{44,57} Interestingly, the threefold twins of ordered 4,8-**scu** structure led to a twinned 4,12-**c** rare **shp** structure (Fig. S17, ESI†). To the best of our knowledge, the inherent correlation between these two topological nets has not been realized until this work. As a result, the average structure is observed as triangular-shaped instead of the expected diamond-shaped channels in NU-903 (Fig. 2).

The MOF synthesis with linker L3 yields hexagonal rod-shaped crystals (Fig. 3b). Single crystal X-ray diffraction analysis (Table S1, ESI†) revealed that NU-1008 crystallized in the hexagonal *P6/mmm* space group with the chemical formula as $\text{Zr}_6(\mu\text{-O})_4(\mu\text{-OH})_4(\text{HCOO})(\text{H}_2\text{O})_3(\text{OH})_3(\text{L3})_2$; the 3D structure consists of 8-connected Zr₆ nodes and dibromo-substituted TCPB linkers in a **csq** topology. Isorecticular to NU-1000,⁵⁸ there are two types of channels along the *c* axis, a 1 nm wide triangular channel and a 3 nm hexagonal channel (Fig. 1). The incorporation of the dibromo groups does not compromise the porosity of the material because the bromo groups reside in the window connecting the hierarchically triangular micropores and hexagonal mesopores. Therefore, the pore size distribution of NU-1008 is similar to NU-1000.

Characterization of as-synthesized MOFs

Due to the 8-connected Zr₆ nodes present in all three MOFs, ¹H NMR was employed to identify the coordinated species on the remaining accessible sites of the node. Upon digestion of the MOFs with a dilute NaOD solution, a peak was observed for all three materials around 8.37 ppm that is characteristic of

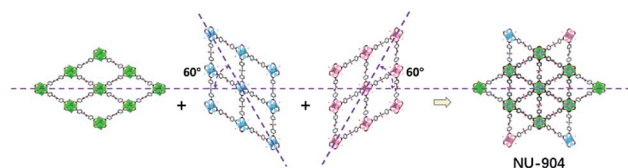


Fig. 2 The reticular twin structure of NU-904.





Fig. 3 Characterization of the three MOFs. (a) PXRD patterns; (b) optical images of the single crystals of NU-904 (left) and NU-1008 (right), the scale bars in the images are 100 μm ; (c) N_2 sorption isotherms and DFT pore size distribution of NU-903, NU-904 and NU-1008.

formate ligand. Quantifying the amounts of formate relative to the organic linker in the three samples gives *ca.* 2, 2 and 3 formate groups per Zr_6 node in NU-903 (Fig. S4, ESI †), NU-904 (Fig. S5, ESI †) and NU-1008 (Fig. S6, ESI †), respectively. Diffuse reflectance infrared fourier transform spectroscopy (DRIFTS) also confirms the presence of formate groups with the observed C–H stretching at 2745 cm^{-1} . In addition, N=O stretching (1658 and 1373 cm^{-1}) and C–Br stretching (712 cm^{-1}) in the DRIFTS spectra of NU-904 and NU-1008 confirmed the presence of L2 and L3, respectively (Fig. S10–S12, ESI †).

The phase purity of the bulk materials was confirmed by PXRD (Fig. 3a). The particle morphology of NU-904 is oval-shaped and NU-1008 is hexagonal rod-shaped (Fig. 3b), as evidenced by the SEM images (Fig. S13, ESI †). Nitrogen sorption isotherms of three different MOFs were measured at 77 K; the type I isotherms of NU-903 and NU-904 are indicative of microporous structure of the MOFs, while the type IV isotherm of NU-1008 indicates the presence of both micropores and mesopores in the material. The BET areas were calculated to be 1140, 1410 and $1400\text{ m}^2\text{ g}^{-1}$ for NU-903, NU-904 and NU-1008, respectively. DFT pore size distributions reveal micropores of 11 Å for NU-903, 10 Å for NU-904 and hierarchical micropores of 11 Å and mesopores of 29 Å for NU-1008 (Fig. 3c). The total pore volumes were 0.515 , 0.613 and $0.819\text{ cm}^3\text{ g}^{-1}$ for NU-903, NU-904 and NU-1008, respectively. The larger gravimetric pore volume and the type IV isotherm of NU-1008 is again consistent with a mesoporous structure. Thermogravimetric analysis (TGA)

under air showed no sign of mass loss up to $400\text{ }^\circ\text{C}$ (Fig. S7–S9, ESI †), demonstrating the high thermal stability of the three MOFs. In addition, all three materials showed good stability under acidic condition with treatment of 0.5 M HCl aqueous solution for 10 hours (Fig. S14, ESI †).

Topology and porosity control through introduction of steric effect

To investigate the influence of linker conformation on topology control, we performed a detailed analysis of linker conformation in the three crystal structures. In the *scu*-net NU-903 structure,⁴⁵ the TCPB linker with no substituents adopts a C_{2h} symmetry with the C_2 axis and the perpendicular σ_h depicted in Fig. 4a. In the MOF structure, the two adjacent phenyl arms in the upper side rotate away from each other, whereas the lower ones rotate toward each other (Fig. 4a). In NU-904, the introduction of the mononitro group lowers the symmetry of the linker to C_2 but the conformation of the peripheral phenyl arms was found to be similar to the non-substituted TCPB linker in NU-903 (Fig. 4b). As a result, each of the twinning portions that comprise this structure has the same *scu* topology as NU-903, even though the average structure of NU-904 has 1D triangular channels instead of the diamond-shaped channels with a typical *scu* topology. Notably, the mononitro group on the central benzene rotates $\sim 32^\circ$ to fit in the structure. In NU-1008 with dibromo-substituted TCPB, the linker adopts a C_{2v} symmetry due to the rotation of the upper and lower pair of phenyl arms toward each other (Fig. 4c) which directs the





Fig. 4 Linker conformation in (a) NU-903, (b) NU-904 and (c) NU-1008.

framework to **csq** topology over the **scu**-net MOF. In addition, to accommodate the dibromo groups in NU-1008, the dihedral angle between the arm benzene and central benzene in the **csq**-net NU-1008 is closer to 90° than the dihedral angle observed in **scu**-net NU-903 and NU-904 (Table S3, ESI[†]). Additionally, the angle between the arms (117°) in NU-1008 is larger than in **scu**-net NU-903 (115°) and NU-904 (108.5°).

The 8-connected Zr_6 cluster in the three MOFs can be regarded as a cubic shaped node and the linker as a rectangular plane. In NU-903, the non-substituted TCPB linker with C_{2h} symmetry can link the clusters in the same orientation (Fig. 1), giving the **scu** topology. In contrast, the Zr_6 clusters in NU-1008 undergo an orientation change to conform to the C_{2v} symmetry of dibromo TCPB linker, and form a different topology, **csq**. The topology change confirmed our hypothesis that the steric effect induced by linker functionalization can achieve the topology control of MOF materials.

Catalytic performance for CO₂ fixation under mild conditions

CO₂ is a well-established greenhouse gas that originates from the carbon footprint of human activities.⁵⁹ In order to mitigate the adverse effect of CO₂ to the environment, considerable progress

has been made in its capture and storage in recent years,^{60,61} and promising strategies for the consumption of CO₂ have received much attention. Rational utilization of CO₂ is of great significance for the deceleration of global warming and the development of sustainable energy. Chemical fixation of CO₂ with epoxides through a facile cycloaddition process, catalysed by acidic sites, is an attractive route to convert this greenhouse gas into highly demanded cyclic organic carbonates.⁶²

Owing to the presence of substrate accessible Lewis acidic sites, MOFs have been previously explored to catalyse the cycloaddition of CO₂ and epoxides.^{63,64} However, energy-demanding reaction conditions such as elevated temperatures and pressures are generally required for the efficient conversion of CO₂ into cyclic carbonates.⁶⁵ Inspired by these reports, CO₂ fixation into styrene oxide was used as a model reaction to test the catalytic performance of MOFs reported in this study under mild conditions.

NU-903, NU-904 and NU-1008 were investigated for CO₂ fixation into styrene oxide under room temperature with 1 bar of CO₂. The styrene oxide (0.2 mmol), tetrabutylammonium bromide (6.5 mg, 0.02 mmol) pre-dissolved in 400 μ L of acetonitrile and MOF material (4.0 mol%, 0.002 mmol) were added to an autoclave batch reactor, which had previously been dried for 6 h at 80°C . The autoclave reactor was evacuated, purged with CO₂, and then placed under a constant pressure of CO₂ under 5 bar for 15 min to allow the system to equilibrate. The reaction was carried out at room temperature for 24 hours after the pressure was reduced to 1 bar of gauge pressure. After the reaction, the catalyst was separated by centrifugation and a small aliquot of the supernatant reaction mixture was taken to be analyzed by ¹H NMR to calculate the conversion (Fig. 5).



Fig. 5 Proton NMR spectra of CO₂ fixation into styrene oxide catalyzed by NU-903, NU-904 and NU-1008 (1 bar, room temperature, 24 h; CDCl₃ as deuterium solvent, 500 MHz). The peak with asterisk refers to the tetrabutylammonium bromide.



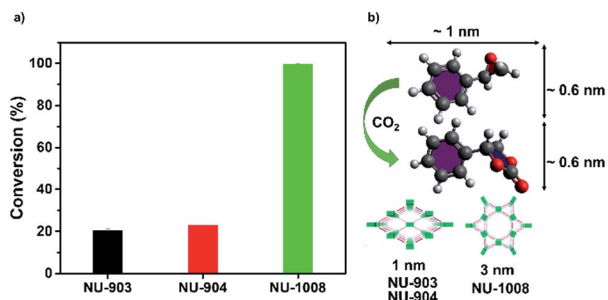


Fig. 6 CO₂ fixation conversion on NU-903, NU-904 and NU-1008. (a) Catalytic capacities; (b) size of substrate, product and pore size of NU-903, NU-904 and NU-1008.

A control experiment without any MOF catalyst exhibited negligible conversion of styrene oxide. However, vastly different CO₂ fixation activities were observed within the three MOF catalysts utilized in this report. Remarkably, full conversion of styrene oxide was observed with NU-1008 after 24 hours under mild conditions (room temperature, 1 bar of CO₂). Previous reports using MOF catalysts for this reaction required elevated temperatures (~120 °C), pressures (10–20 bar) or longer reaction time (up to 56 hours) to afford similar or lower substrate conversions (Table S2, ESI[†]).⁶² Thus, NU-1008 stands out as a more environmentally friendly solid acid catalyst for the chemical fixation of CO₂, compared to others reported.

In comparison, NU-903 and NU-904 showed much lower conversion (Fig. 6a), despite the fact that they had similar CO₂ adsorption performance to NU-1008 (Fig. S15, ESI[†]). This significant difference is likely attributed to the aforementioned narrower pores and smaller pore volumes in NU-903 and NU-904 (around 10 Å, 0.50 cm³ g⁻¹) compared to the mesoporous channels in NU-1008 (around 30 Å, 0.82 cm³ g⁻¹). The microporous NU-903 and NU-904 most likely limited the diffusion of the sizeable substrate and product (Fig. 6b) and resulted in the lower conversion. The turnover numbers (TONs) in 24 hours were calculated to be 20.3, 22.8 and 99.4 for NU-903, NU-904 and NU-1008, respectively.

Conclusions

Three different Zr-MOFs were constructed by the steric control of the conformations of substituted tetatopic carboxylate linkers. Compared to NU-903 with the **scu** topology, phase pure **csq**-net NU-1008 supports our initial hypothesis that introducing functionality into the organic linker can influence the linker conformation and direct the topology of targeted MOFs. Consequently, some mechanistic insights regarding MOF synthesis can be inferred in efforts to make MOF topology control possible with the strategic design of organic linkers. Interestingly, with nitro groups, threefold twinning is observed in **scu**-net NU-904, which yields an average structure of **shp**-net with triangular microchannels while maintaining the unsaturated metal sites for promising catalytic applications. Significantly different CO₂ fixation catalytic activities were observed due to the diverse pore structures. Among them, the

mesoporous NU-1008 is found to be a highly active MOF catalyst which displays complete CO₂ fixation into styrene oxide in less than 24 hours under room temperature and 1 bar of CO₂.

Conflicts of interest

There are no conflicts to declare.

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Notes and references

- O. M. Yaghi, M. O'Keeffe, N. W. Ockwig, H. K. Chae, M. Eddaoudi and J. Kim, *Nature*, 2003, **423**, 705.
- A. I. Cooper and M. J. Rosseinsky, *Nat. Chem.*, 2009, **1**, 26.
- G. K. H. Shimizu, R. Vaidhyanathan and J. M. Taylor, *Chem. Soc. Rev.*, 2009, **38**, 1430–1449.
- P. Nugent, Y. Belmabkhout, S. D. Burd, A. J. Cairns, R. Luebke, K. Forrest, T. Pham, S. Ma, B. Space, L. Wojtas, M. Eddaoudi and M. J. Zaworotko, *Nature*, 2013, **495**, 80.
- K. Sumida, D. L. Rogow, J. A. Mason, T. M. McDonald, E. D. Bloch, Z. R. Herm, T.-H. Bae and J. R. Long, *Chem. Rev.*, 2012, **112**, 724–781.
- X. Cui, K. Chen, H. Xing, Q. Yang, R. Krishna, Z. Bao, H. Wu, W. Zhou, X. Dong, Y. Han, B. Li, Q. Ren, M. J. Zaworotko and B. Chen, *Science*, 2016, **353**, 141–144.
- D.-X. Xue, A. J. Cairns, Y. Belmabkhout, L. Wojtas, Y. Liu, M. H. Alkordi and M. Eddaoudi, *J. Am. Chem. Soc.*, 2013, **135**, 7660–7667.
- J.-J. Zheng, S. Kusaka, R. Matsuda, S. Kitagawa and S. Sakaki, *J. Phys. Chem. C*, 2017, **121**, 19129–19139.
- E. J. Cussen, J. B. Claridge, M. J. Rosseinsky and C. J. Kepert, *J. Am. Chem. Soc.*, 2002, **124**, 9574–9581.
- L. Li, R.-B. Lin, R. Krishna, X. Wang, B. Li, H. Wu, J. Li, W. Zhou and B. Chen, *J. Am. Chem. Soc.*, 2017, **139**, 7733–7736.
- A. J. Rieth and M. Dincă, *J. Am. Chem. Soc.*, 2018, **140**, 3461–3466.



- 12 X. Zhang, Z. Huang, M. Ferrandon, D. Yang, L. Robison, P. Li, T. C. Wang, M. Delferro and O. K. Farha, *Nature Catalysis*, 2018, **1**, 356–362.
- 13 Z. Li, A. W. Peters, V. Bernales, M. A. Ortuño, N. M. Schweitzer, M. R. DeStefano, L. C. Gallington, A. E. Platero-Prats, K. W. Chapman, C. J. Cramer, L. Gagliardi, J. T. Hupp and O. K. Farha, *ACS Cent. Sci.*, 2017, **3**, 31–38.
- 14 C. R. Kim, T. Uemura and S. Kitagawa, *Chem. Soc. Rev.*, 2016, **45**, 3828–3845.
- 15 H. D. Park, M. Dincă and Y. Román-Leshkov, *J. Am. Chem. Soc.*, 2018, **140**, 10669–10672.
- 16 S. M. J. Rogge, A. Bavykina, J. Hajek, H. Garcia, A. I. Olivos-Suarez, A. Sepúlveda-Escribano, A. Vimont, G. Clet, P. Bazin, F. Kapteijn, M. Daturi, E. V. Ramos-Fernandez, F. X. Llabrés i Xamena, V. Van Speybroeck and J. Gascon, *Chem. Soc. Rev.*, 2017, **46**, 3134–3184.
- 17 D. Aulakh, J. B. Pyser, X. Zhang, A. A. Yakovenko, K. R. Dunbar and M. Wriedt, *J. Am. Chem. Soc.*, 2015, **137**, 9254–9257.
- 18 X. Zhang, M. R. Saber, A. P. Prosvirin, J. H. Reibenspies, L. Sun, M. Ballesteros-Rivas, H. Zhao and K. R. Dunbar, *Inorg. Chem. Front.*, 2015, **2**, 904–911.
- 19 Y. Chen, P. Li, J. A. Modica, R. J. Drout and O. K. Farha, *J. Am. Chem. Soc.*, 2018, **140**, 5678–5681.
- 20 R. Freund, U. Lächelt, T. Gruber, B. Rühle and S. Wuttke, *ACS Nano*, 2018, **12**, 2094–2105.
- 21 S. Wuttke, M. Lismont, A. Escudero, B. Rungtaweeworanit and W. J. Parak, *Biomaterials*, 2017, **123**, 172–183.
- 22 S. Wang, C. M. McGuirk, A. d'Aquino, J. A. Mason and C. A. Mirkin, *Adv. Mater.*, 2018, **30**, 1800202.
- 23 T. Simon-Yarza, A. Mielcarek, P. Couvreur and C. Serre, *Adv. Mater.*, 2018, **30**, 1707365.
- 24 B. Chen, S. Xiang and G. Qian, *Acc. Chem. Res.*, 2010, **43**, 1115–1124.
- 25 Z. Hu, B. J. Deibert and J. Li, *Chem. Soc. Rev.*, 2014, **43**, 5815–5840.
- 26 X. Zhang, Q. Hu, T. Xia, J. Zhang, Y. Yang, Y. Cui, B. Chen and G. Qian, *ACS Appl. Mater. Interfaces*, 2016, **8**, 32259–32265.
- 27 H. Furukawa, F. Gándara, Y.-B. Zhang, J. Jiang, W. L. Queen, M. R. Hudson and O. M. Yaghi, *J. Am. Chem. Soc.*, 2014, **136**, 4369–4381.
- 28 J. H. Cavka, S. Jakobsen, U. Olsbye, N. Guillou, C. Lamberti, S. Bordiga and K. P. Lillerud, *J. Am. Chem. Soc.*, 2008, **130**, 13850–13851.
- 29 J. Lippke, B. Brosent, T. von Zons, E. Virmani, S. Lilienthal, T. Preuß, M. Hülsmann, A. M. Schneider, S. Wuttke, P. Behrens and A. Godt, *Inorg. Chem.*, 2017, **56**, 748–761.
- 30 V. Guillerm, Ł. J. Weseliński, Y. Belmabkhout, A. J. Cairns, V. D'Elia, Ł. Wojtas, K. Adil and M. Eddaoudi, *Nat. Chem.*, 2014, **6**, 673.
- 31 Z. Chen, Ł. J. Weseliński, K. Adil, Y. Belmabkhout, A. Shkurenko, H. Jiang, P. M. Bhatt, V. Guillerm, E. Dauzon, D.-X. Xue, M. O'Keeffe and M. Eddaoudi, *J. Am. Chem. Soc.*, 2017, **139**, 3265–3274.
- 32 C.-W. Kung, T. C. Wang, J. E. Mondloch, D. Fairen-Jimenez, D. M. Gardner, W. Bury, J. M. Klingsporn, J. C. Barnes, R. Van Duyn, J. F. Stoddart, M. R. Wasielewski, O. K. Farha and J. T. Hupp, *Chem. Mater.*, 2013, **25**, 5012–5017.
- 33 H. Noh, C.-W. Kung, T. Islamoglu, A. W. Peters, Y. Liao, P. Li, S. J. Garibay, X. Zhang, M. R. DeStefano, J. T. Hupp and O. K. Farha, *Chem. Mater.*, 2018, **30**, 2193–2197.
- 34 J. Ma, L. D. Tran and A. J. Matzger, *Cryst. Growth Des.*, 2016, **16**, 4148–4153.
- 35 D. Feng, Z.-Y. Gu, J.-R. Li, H.-L. Jiang, Z. Wei and H.-C. Zhou, *Angew. Chem., Int. Ed.*, 2012, **51**, 10307–10310.
- 36 W. Morris, B. Voloskiy, S. Demir, F. Gándara, P. L. McGrier, H. Furukawa, D. Cascio, J. F. Stoddart and O. M. Yaghi, *Inorg. Chem.*, 2012, **51**, 6443–6445.
- 37 Y. Chen, T. Hoang and S. Ma, *Inorg. Chem.*, 2012, **51**, 12600–12602.
- 38 S. B. Kalidindi, S. Nayak, M. E. Briggs, S. Jansat, A. P. Katsoulidis, G. J. Miller, J. E. Warren, D. Antypov, F. Corà, B. Slater, M. R. Prestly, C. Marti-Gastaldo and M. J. Rosseinsky, *Angew. Chem., Int. Ed.*, 2015, **54**, 221–226.
- 39 T. C. Wang, W. Bury, D. A. Gómez-Gualdrón, N. A. Vermeulen, J. E. Mondloch, P. Deria, K. Zhang, P. Z. Moghadam, A. A. Sarjeant, R. Q. Snurr, J. F. Stoddart, J. T. Hupp and O. K. Farha, *J. Am. Chem. Soc.*, 2015, **137**, 3585–3591.
- 40 D. Feng, H.-L. Jiang, Y.-P. Chen, Z.-Y. Gu, Z. Wei and H.-C. Zhou, *Inorg. Chem.*, 2013, **52**, 12661–12667.
- 41 Q. Lin, X. Bu, A. Kong, C. Mao, X. Zhao, F. Bu and P. Feng, *J. Am. Chem. Soc.*, 2015, **137**, 2235–2238.
- 42 J. E. Mondloch, W. Bury, D. Fairen-Jimenez, S. Kwon, E. J. DeMarco, M. H. Weston, A. A. Sarjeant, S. T. Nguyen, P. C. Stair, R. Q. Snurr, O. K. Farha and J. T. Hupp, *J. Am. Chem. Soc.*, 2013, **135**, 10294–10297.
- 43 D. Feng, W.-C. Chung, Z. Wei, Z.-Y. Gu, H.-L. Jiang, Y.-P. Chen, D. J. Darensbourg and H.-C. Zhou, *J. Am. Chem. Soc.*, 2013, **135**, 17105–17110.
- 44 D. Feng, Z.-Y. Gu, Y.-P. Chen, J. Park, Z. Wei, Y. Sun, M. Bosch, S. Yuan and H.-C. Zhou, *J. Am. Chem. Soc.*, 2014, **136**, 17714–17717.
- 45 M. Lammert, H. Reinsch, C. A. Murray, M. T. Wharmby, H. Terraschke and N. Stock, *Dalton Trans.*, 2016, **45**, 18822.
- 46 S. Wang, J. Wang, W. Cheng, X. Yang, Z. Zhang, Y. Xu, H. Liu, Y. Wu and M. Fang, *Dalton Trans.*, 2015, **44**, 8049–8061.
- 47 M. Zhang, Y.-P. Chen, M. Bosch, T. Gentle, K. Wang, D. Feng, Z. U. Wang and H.-C. Zhou, *Angew. Chem., Int. Ed.*, 2014, **53**, 815–818.
- 48 H.-L. Jiang, D. Feng, K. Wang, Z.-Y. Gu, Z. Wei, Y.-P. Chen and H.-C. Zhou, *J. Am. Chem. Soc.*, 2013, **135**, 13934–13938.
- 49 Y. Zhang, X. Zhang, J. Lyu, K. I. Otake, X. Wang, L. R. Redfern, C. D. Malliakas, Z. Li, T. Islamoglu, B. Wang and O. K. Farha, *J. Am. Chem. Soc.*, 2018, **140**, 11179–11183.
- 50 P. T. K. Nguyen, H. T. D. Nguyen, H. N. Nguyen, C. A. Trickett, Q. T. Ton, E. Gutiérrez-Puebla, M. A. Monge, K. E. Cordova and F. Gándara, *ACS Appl. Mater. Interfaces*, 2018, **10**, 733–744.



- 51 T. Islamoglu, K.-i. Otake, P. Li, C. T. Buru, A. W. Peters, I. Akpınar, S. J. Garibay and O. K. Farha, *CrystEngComm*, 2018, **20**, 5913–5918.
- 52 P. Li, Q. Chen, T. C. Wang, N. A. Vermeulen, B. L. Mehdi, A. Dohnalkova, N. D. Browning, D. Shen, R. Anderson, D. A. Gómez-Gualdrón, F. M. Cetin, J. Jagiello, A. M. Asiri, J. F. Stoddart and O. K. Farha, *Chemistry*, 2018, **4**, 1022–1034.
- 53 Y.-B. Zhang, H. Furukawa, N. Ko, W. Nie, H. J. Park, S. Okajima, K. E. Cordova, H. Deng, J. Kim and O. M. Yaghi, *J. Am. Chem. Soc.*, 2015, **137**, 2641–2650.
- 54 J. Pang, S. Yuan, J. Qin, C. Liu, C. Lollar, M. Wu, D. Yuan, H.-C. Zhou and M. Hong, *J. Am. Chem. Soc.*, 2017, **139**, 16939–16945.
- 55 W.-G. Liu and D. Truhlar, *Chem. Mater.*, 2017, **29**, 8073–8081.
- 56 P. Deria, D. A. Gómez-Gualdrón, I. Hod, R. Q. Snurr, J. T. Hupp and O. K. Farha, *J. Am. Chem. Soc.*, 2016, **138**, 14449–14457.
- 57 R. G. AbdulHalim, P. M. Bhatt, Y. Belmabkhout, A. Shkurenko, K. Adil, L. J. Barbour and M. Eddaoudi, *J. Am. Chem. Soc.*, 2017, **139**, 10715–10722.
- 58 P. Deria, J. E. Mondloch, E. Tylianakis, P. Ghosh, W. Bury, R. Q. Snurr, J. T. Hupp and O. K. Farha, *J. Am. Chem. Soc.*, 2013, **135**, 16801–16804.
- 59 S. Solomon, G.-K. Plattner, R. Knutti and P. Friedlingstein, *Proc. Natl. Acad. Sci. U. S. A.*, 2009, **106**, 1704–1709.
- 60 C. A. Trickett, A. Helal, B. A. Al-Maythaly, Z. H. Yamani, K. E. Cordova and O. M. Yaghi, *Nat. Rev. Mater.*, 2017, **2**, 17045.
- 61 J. Yu, L.-H. Xie, J.-R. Li, Y. Ma, J. M. Seminario and P. B. Balbuena, *Chem. Rev.*, 2017, **117**, 9674–9754.
- 62 M. H. Beyzavi, C. J. Stephenson, Y. Liu, O. Karagiari, J. T. Hupp and O. K. Farha, *Front. Energy Res.*, 2015, **2**, 63.
- 63 J. Kim, S.-N. Kim, H.-G. Jang, G. Seo and W.-S. Ahn, *Appl. Catal., A*, 2013, **453**, 175–180.
- 64 M. H. Beyzavi, R. C. Klet, S. Tussupbayev, J. Borycz, N. A. Vermeulen, C. J. Cramer, J. F. Stoddart, J. T. Hupp and O. K. Farha, *J. Am. Chem. Soc.*, 2014, **136**, 15861–15864.
- 65 C. Martín, G. Fiorani and A. W. Kleij, *ACS Catal.*, 2015, **5**, 1353–1370.

