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Charge transfer excitons in unfunctionalized graphite-wrapped MAPbBr₃ nanocrystal composites with different morphologies†

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Highly luminescent methylammonium lead halide perovskite nanocrystals (PNCs) have attracted considerable research interest due to their alluring properties that can be tuned as per the morphological variations. Charge transfer excitons (CTEs) are a good representation of the capacity of semiconductor materials to tune their properties. We report here that CTEs may be generated successfully in graphite:PNC composites functioning as host–guest components. These hybrid perovskite nanocrystals are embedded in graphite sheets that are unfunctionalized. The PL spectrum demonstrated that graphite significantly reduces the luminescence intensity of the perovskite nanocrystals. The morphologies of the perovskite nanocrystals are developed in such a manner that they may be readily incorporated onto the surface of graphite sheets to facilitate effective charge transfer.

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Introduction

Hybrid lead halide perovskite nanocrystals have been identified as a promising alternative for semiconducting materials in the optoelectronics field due to their outstanding properties, such as high PLQY, narrowband emission, high tunability, long carrier diffusion length, and tunable absorption and emission.^{1–9} The most significant accomplishment of these materials is the rapid increase in power conversion efficiency (PCE) in solar cell applications, from 3.8% to 25.5%.^{10–13} Perovskite materials will play an important part in the development of solar energy in the future since solar energy is a key component of the world's most important and limitless source of energy. As a result of the quantum confinement effect, the properties of these materials may be tuned based on the size and shape of the nanocrystals. The CTEs, on the other hand, are artificially engineered excitons that have been introduced in order to produce high-efficiency optoelectronic functions.^{14,15} CTEs may be created in heterostructures when charge transfer between distinct local structures results in the formation of electron–hole pairs. Heterostructures are generated when distinct donor acceptor components are mixed. Additionally, CTEs

have been discovered between 2D perovskites and organic molecules where charge transfer is facilitated by orbital overlap.¹⁶ Perovskites are utilized in optoelectronic devices as an active layer or a transport layer (either hole or electron).

Carbon compounds with adequate electrical and electronic characteristics are employed as electrodes in LEDs and solar cells. The efficiency of the device is governed by the degree to which the charge-collection and charge-injection characteristics of the materials are affected.^{17–23} Carbon allotropes such as carbon nanotubes, graphene, graphene oxide, and graphite are utilized in combination with metal halide perovskites to perform a variety of functions.²⁴ Petra J. Cameron and coworkers demonstrated the stability of lead halide perovskite (CsPbBr₃) encapsulated in graphite sheets and mesoporous carbon scaffolds. Not only do they exhibit stability, but they also illustrate the employment of photoanodes as a carrier and protective layer.²⁵ Surendra K. Yadav employed a compressed carbon/graphite nanocomposite as a hydrophobic cathode in a metal halide perovskite.²⁶ Andreas Hinsch and his group investigated the influence of natural and grown graphite particles on the optoelectronic characteristics of carbon-based perovskite cells (c-PSCs).²⁷ Apart from this, carbon derivatives are extensively employed for charge transfer owing to their adequate energy levels, low cost, and abundance on the planet.^{28–30}

Mangalaraja *et al.*³¹ reported the charge transfer of graphene with the perovskite materials through the electro-optic (E-O) switching modulation. Time-resolved luminescence and dielectric spectroscopy were used to study the carrier dynamics and transfer mechanisms in the nanocomposite. Loh *et al.*³² reported the charge

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transfer and van der Waals interaction between graphene and Ruddlesden Popper perovskite. Zhang *et al.* reported the charge transfer of CsPbBr₃ nanowires with MoS₂.³³ Yu and his group reported the charge transfer of layered perovskites with MoS₂ through the chemical vapor deposition technique.³⁴ Gan *et al.* reported the charge transfer of CsPb(Br_xI_{1-x})₃ nanocrystals³⁵ with C₄N₃, which was confirmed through temperature-dependent PL and time-resolved PL studies.

Motivated by the remarkable progress made in optoelectronic applications,^{36–39} our group has recently reported the use of perovskite nanocrystals and carbon nanotube composites for efficient charge transfer studies.⁴⁰ Apart from the aforementioned specific system, semiconductor nanomaterials are used in photoelectrodes to conduct charge transfer experiments, which have a plethora of appealing features.^{25,26}

In this work, we primarily employed the chemical technique to modify the emission range of MAPbBr₃ by monitoring the ratio of used passivated ligands and preparing MAPbBr₃ perovskite nanostructures of diverse morphologies. The heterostructures are prepared by mixing these perovskite nanostructures with graphite sheets. The wide emission band with an extended lifetime is the first evidence of the creation of CTEs at the interfaces. The investigation has been done using steady-state photoluminescence, time-correlated single-photon counting (TCSPC), and absolute photoluminescence quantum yield (PLQY). This work presents a simple yet efficient approach for charge transfer investigations in halide hybrid perovskite materials along with focusing attention on its prospective applications in the future.

Experimental section

Materials and methods

All the chemicals were purchased from commercial sources and were used as received without further purification. Lead(II) bromide (99%), methylamine solution (33 wt% in absolute ethanol), oleic acid (OA), and oleyl amine (OAm) were purchased from Sigma-Aldrich. Other reagents like DMF (SRL),

ethanol, hydrobromic acid (Rankem) (48 wt% in water), and graphite synthetic (type 2) (SRL) were obtained from a local vendor.

Methylammonium bromide (MABr) synthesis. Briefly, in a 50 mL flask, 5.14 mL of CH₃NH₂ was mixed with 10 mL of ethanol (33 wt% in absolute ethanol) at 0 °C under vigorous stirring. Then, 6 mL of HBr was added dropwise with continuous stirring at room temperature. The obtained solution was placed in a rotary evaporator at 60 °C to remove all the solvents. After washing many times with diethyl ether, a solid product of MABr was obtained.

Synthesis of MAPbBr₃. The ligand-assisted reprecipitation method was used for the synthesis of perovskite nanocrystals. 0.1 mmol of methylammonium bromide (MABr) and 0.1 mmol of PbBr₂ were dissolved in 1 mL of DMF with 100 µL of oleic acid and a varying amount of oleylamine from 35 to 300 µL. At the end of the reaction, a clear solution was formed. Then 30 µL of the precursor was injected into 3 mL of chloroform to obtain the desired nanocrystals.

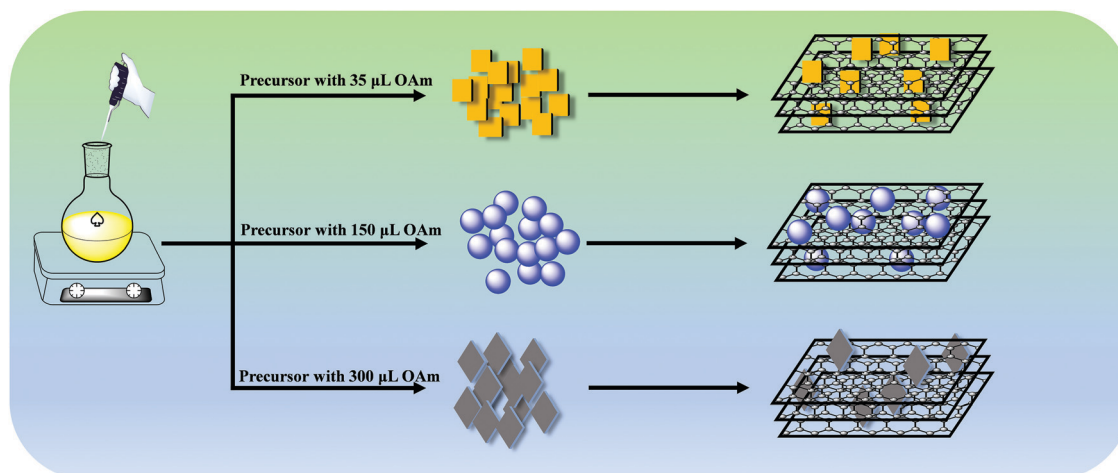
Synthesis of P35 and P35@G. 11.1 mg of MABr and 36.7 mg of PbBr₂ were dissolved in 1 mL of DMF with 100 µL of oleic acid and 35 µL of oleylamine for **P35**. Then, 30 µL of precursor solution was added to 3 mL of chloroform. This solution was centrifuged at 7000 rpm speed for 10 minutes and again redispersed in chloroform, which was further used for characterization.

60 mg of graphite was dissolved in 10 mL of chloroform (0.6% w/v in chloroform) to synthesize the perovskite NC/graphite nanocomposite. It was sonicated for 5 mins so that graphite can disperse in chloroform. Then 30 µL of the precursor **P35** was mixed with the graphite solution to form the desired nanocomposite and named **P35@G**.

Synthesis details of other samples have been incorporated into the ESI.†

Results and discussion

MAPbBr₃ NCs are synthesized by using the procedure developed in our laboratory. Briefly, the precursor complex is



Scheme 1 Synthesis of hybrid lead halide perovskite NC and the perovskite NC/graphite nanocomposite.

synthesized by taking a 1 : 1 ratio of MABr and PbBr₂ dissolved in DMF. While the oleic acid is kept constant, the amount of oleyl amine is increased gradually to get different morphologies. The effect of ligand concentration plays a vital role in the growth of perovskite particles. As the concentration of the ligand increases the passivation takes place very effectively on the precursors and slows the growth rate of perovskite particles. The growth of nanoparticles slows down with the increase of the amount of oleyl amine and the particle size decreases, which affects the optical and morphological properties of the PNCs.^{41–43} Various morphologies of PNCs are obtained using different ratios of oleic acid and oleyl amine, as shown in Scheme 1. The synthesized perovskite nanocrystals are integrated into unfunctionalized graphite sheets for the preparation of a nanocomposite. The optical absorption studies revealed that the PNC has a substantial effect on the absorption characteristics of the nanocomposite structures as shown in Fig. S1 (ESI†). Bare perovskite NCs synthesized using 35–150 μ L of oleyl amine exhibit a strong excitonic band at around 400–480 nm. The excitonic band is shifted to a higher wavelength after graphite integration. The same excitonic band appears around 400–450 nm with a minute redshift upon addition of graphite when 200–300 μ L of oleyl amine is used. The redshift of the excitonic band of the perovskite NCs is attributed to the efficient composite formation. The emission spectra of **P35** to **P300** are shown in Fig. 1a, and daylight and UV light images of **P35** to **P300** are shown in Fig. 1b and c. Interestingly, a broad emission is observed between 440 and 510 nm, while the peak position is located at 451, 449, and 441 nm for the composites of **P200**, **P250**, and **P300**, respectively. **P35**, **P100**, and **P150** show a redshift of the luminescence band, which is in accordance with their morphologies.

To obtain more insight into the structure of the crystalline materials and the graphite composite, a concentrated solution of MAPbBr₃ is deposited on a glass substrate and X-ray diffraction patterns were measured (Fig. 2). The diffraction peaks of MAPbBr₃ perovskite nanocrystals observed at 2θ values of 15.38°, 21.63°, 30.61°, 34.25°, 43.60°, and 46.41° correspond to the (010), (011), (020), (021), (022), and (122) planes. The X-ray diffraction pattern shows the pure cubic face of MAPbBr₃

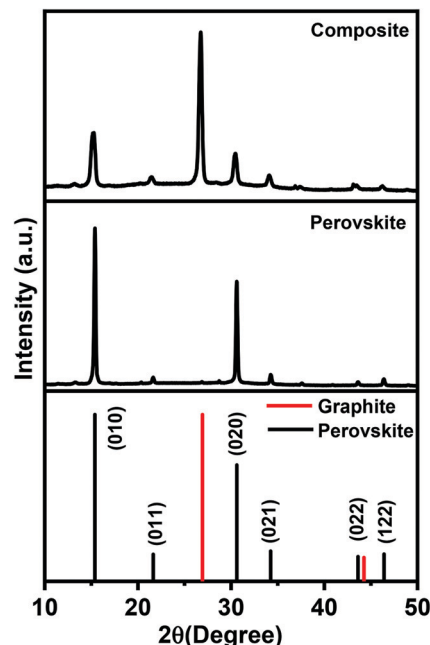


Fig. 2 XRD spectra of pristine perovskite and the PNCs/graphite composite.

(COD no. 1545320) with the $Pm\bar{3}m$ space group.⁴⁴ The crystal planes of the PNCs/graphite composite shifted towards a low theta value compared to those of pristine MAPbBr₃, and the FWHM values of the composite peaks also increase. Enlarged areas of the peaks around $\sim 15^\circ$ and $\sim 30^\circ$ are shown in Fig. S2 in the ESI†. Although the change in theta value and FWHM is significantly very small, it shows the interaction of graphite with MAPbBr₃. Details of peak shifting and change in FWHM value are shown in a tabular form in the ESI† in Table S1. The intense peak in Fig. 2 around 26.7° corresponding to the (003) plane in the composite spectrum shows the presence of graphite (COD no. 1200018).⁴⁵

Scanning electron microscopy (SEM) and transmission electron microscopy (TEM) were performed to understand the morphology of these nanocrystals and the nanocomposite morphological properties.

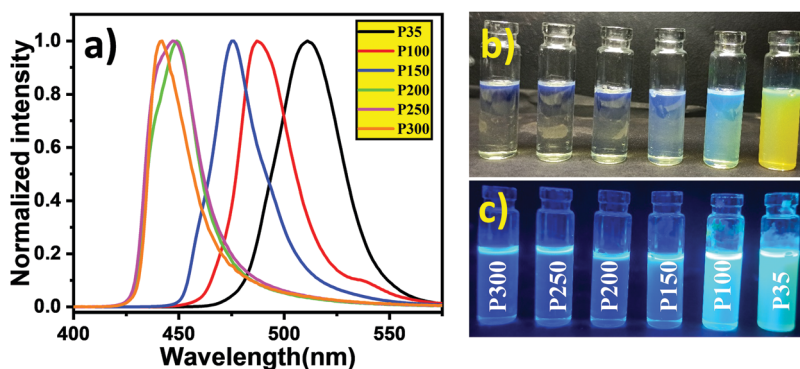


Fig. 1 (a) The emission spectra with tunability of **P35** to **P300** ($\lambda_{\text{exc}} = 415$ nm). Images of **P35** to **P300** samples (b) under visible light and (c) under UV (365 nm) light radiation.

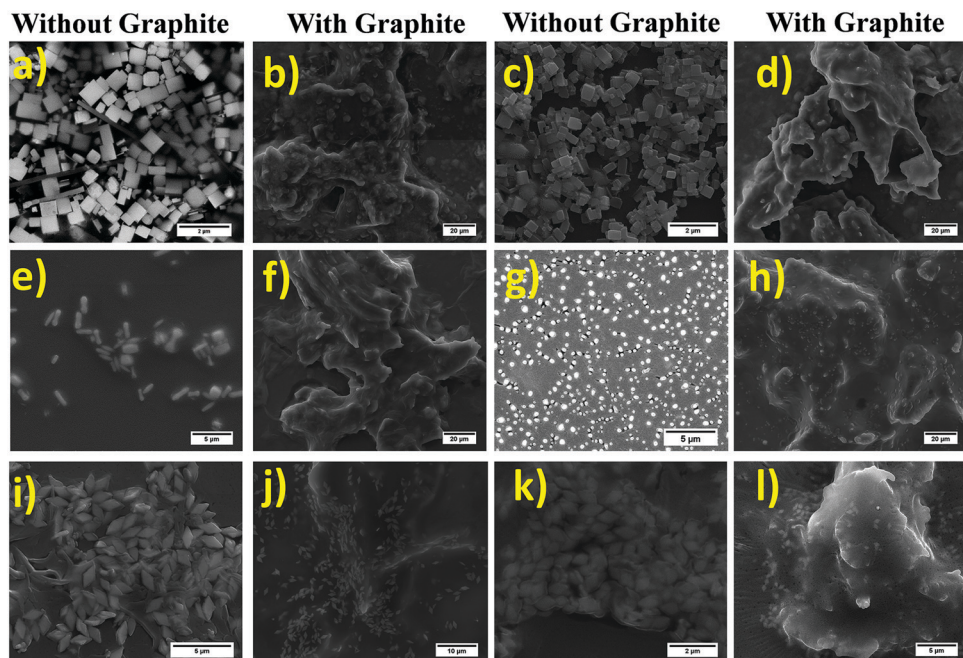


Fig. 3 SEM images of pristine perovskites and corresponding graphite-perovskite composites: (a) **P35**, (b) **P35@G**, (c) **P100**, (d) **P100@G**, (e) **P150**, (f) **P150@G**, (g) **P200**, (h) **P200@G**, (i) **P250**, (j) **P250@G**, (k) **P300**, and (l) **P300@G**.

Firstly, the morphology of the MAPbBr_3 and its evolution with the ligand ratio was investigated by FE-SEM as shown in Fig. 3a-l, which reveals that the nanocrystals are successfully fabricated onto the surface of graphite sheets. The morphology of the perovskite particles was controlled by changing the

ligand concentrations in the reaction medium. SEM images of the perovskite and perovskite-graphite composite at different scale bars are shown in Fig. S3.1–S3.5 (ESI†). The SEM image of graphite is also incorporated in Fig. S3.6 (ESI†). The particles transformed from cubes to rhombus-shape as the

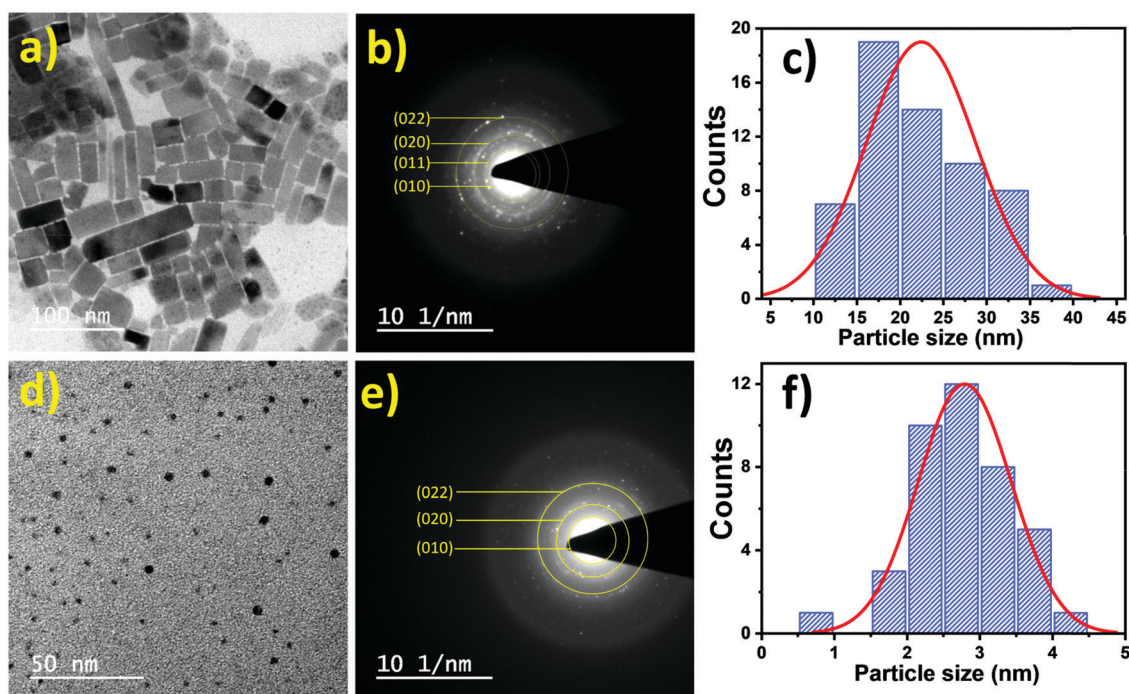


Fig. 4 (a) TEM image, (b) SAED pattern and (c) size distribution curve of **P35**. (d) TEM image, (e) SAED pattern and (f) size distribution curve of **P300**.

concentration of oleyl amine was increased. The self-aggregation phenomenon is taken into account, while considering the development of nanoparticles.

To have a better understanding of the graphite-wrapped perovskite nanocrystals, cross-sectional images of the perovskites and the perovskite-graphite composites are displayed in Fig. S4.1–S4.3 (ESI†). The cross-sectional images demonstrate that the nanocrystals are successfully wrapped with the graphite sheets. TEM studies further reveal that at a lower concentration of ligands, the seeding of nanoparticles occurs. Then, overgrowth of these crystal seeds produces different self-assembled 2D nanoplates depending upon the ligand concentration, as shown in Fig. 4a. On the other hand, a higher concentration of ligand results in quantum dots of less than 5 nm (Fig. 4d). The selected area electron diffraction (SAED) pattern of the nanoparticles exhibits high crystallinity of the materials, as presented in Fig. 4b and e. Simultaneously, we have calculated the corresponding planes obtained from the SAED patterns. Planes calculated from SAED resemble the corresponding XRD planes. Fig. 4c and f show the particle size distribution of 10–40 nm, and 0.5–4.5 nm with the average particle size of ~ 22 nm and ~ 3 nm for the **P35** and **P300** samples, respectively.

X-ray photoelectron spectroscopy (XPS) was further used to investigate the binding energy changes of the Br and Pb atomic environments. A survey scan (Fig. S5, ESI†) confirms the presence of all elements including C, N, Pb, and Br in the pristine and composite perovskite. The Pb and Br peaks of the composite are observed to shift toward higher binding energy. Fig. 5a displays the peaks corresponding to Br $3d_{5/2}$ and Br $3d_{3/2}$ with binding energies at 68.3 and 69.3 eV, respectively, for the pristine perovskite, while Br $3d_{5/2}$ and Br

$3d_{3/2}$ in the composite have values at 68.6, and 69.6 eV. Similarly, Fig. 5b reveals the peaks corresponding to Pb $4f_{5/2}$ and Pb $4f_{7/2}$ with binding energies at 143.3 and 138.4 eV, respectively, for the pristine perovskite. On the other hand, the composite with graphite has values at 143.6, and 138.8 eV for Pb $4f_{5/2}$ and Pb $4f_{7/2}$, respectively. The change in binding energy corroborates the interaction between the perovskite and the composite.

We performed a TGA experiment on PNC and its corresponding composite part. Fig. S6 (ESI†) displays the TGA spectra with a decrease in relative mass with temperature. The first mass-loss step starts from 280 °C to 380 °C where the mass percentage decreased from 100 to 76%. This step corresponds to the degradation of MABr (molar mass: $\text{MABr} = 111.97 \text{ g mol}^{-1}$). In the second step, mass loss starts from 470 °C, and this mass-loss step corresponds to PbBr_2 ($\text{PbBr}_2 = 367.05 \text{ g mol}^{-1}$). Similarly, when we performed TGA for the composite part degradation takes place in a similar manner to that of the pristine perovskite. Here we can conclude that before 250 °C, there is no mass loss in either the pristine perovskite or in the composite part, which is in good accordance with the literature.⁴⁶

Furthermore, we have calculated the absolute PLQY values of all the samples with varying amounts of graphite solution (detail in Fig. S7, and Table S2, ESI†). The PLQY values are 51, 53, 29, 26, 16, and 5 for **P35**, **P100**, **P150**, **P200**, **P250**, and **P300**, respectively. As we increased the amount of graphite solution the PLQY value decreased for each sample.

The emission of PNC is quenched after the formation of nanocomposites. The broad emission band along with the quenching of this emission in the graphite composite leads to the hypothesis that the CTEs are formed at the heterostructure interfaces (Scheme S1, ESI†). To further understand the

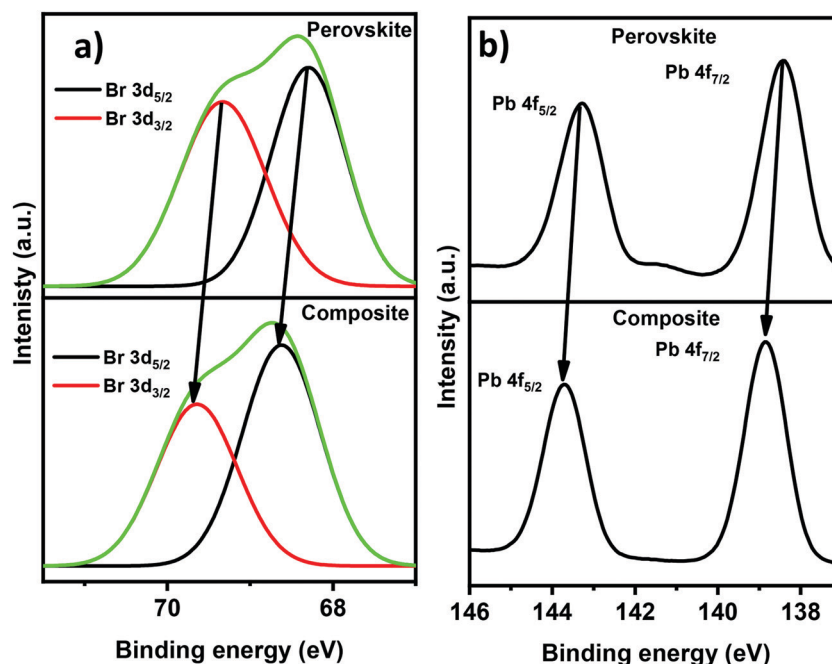


Fig. 5 XPS narrow scans of (a) Br 3d and (b) Pb 4f.

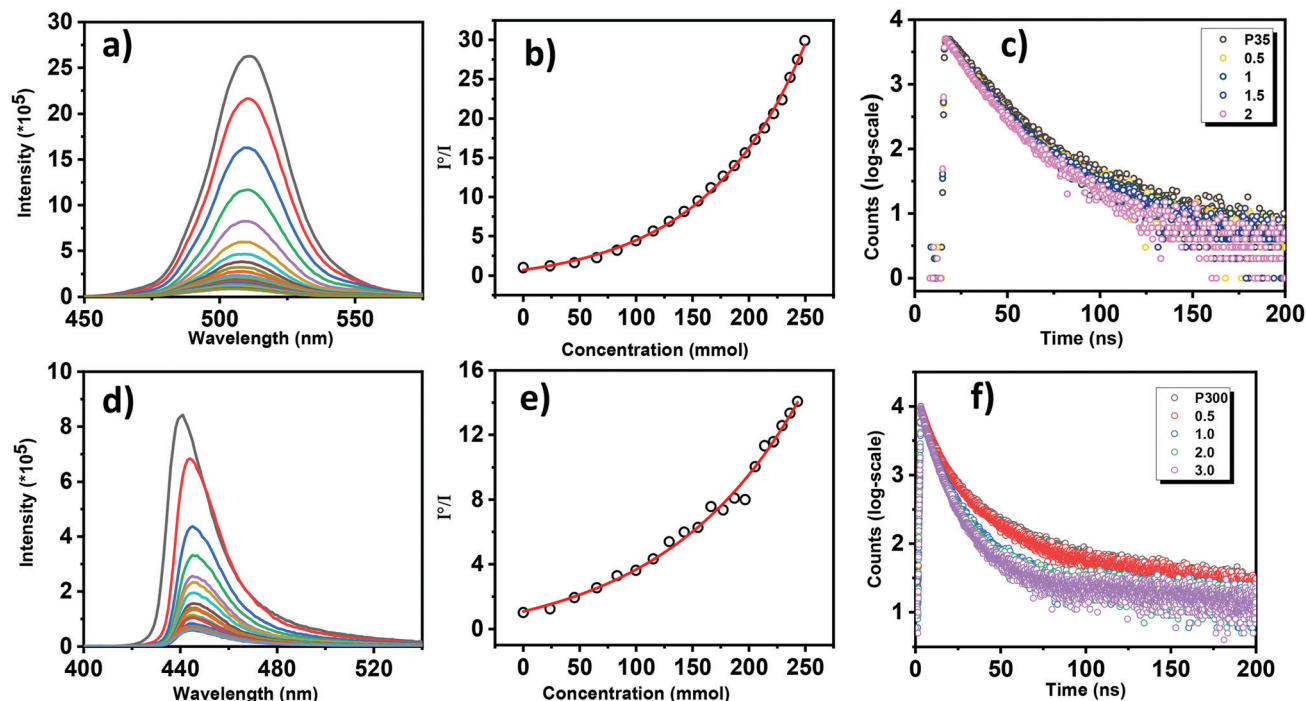


Fig. 6 For **P35**: (a) photoluminescence (PL) quenching with graphite, $\lambda_{\text{exc}} = 415$ nm. (b) Change in PL intensity with respect to the concentration of graphite. (c) TCSPC (laser used = 405 nm). For **P300**: (d) photoluminescence (PL) quenching with graphite, $\lambda_{\text{exc}} = 415$ nm. (e) Change in PL intensity with respect to the concentration of graphite. (f) TCSPC (laser used = 405 nm).

formation of CTEs and the charge transfer, we have performed fluorescence titrations of **P35** to **P300** in which a graphite solution (0.6% w/v in chloroform) was added gradually to the nanocrystals, which are shown in Fig. 6 and Fig. S8 (ESI†).

A continuous decrease in the emission evidenced the charge transfers from **P35** and **P300** to graphite, as noticed in Fig. 6a and d. Similarly, Fig. S7(a–d) (ESI†) correspond to the fluorescence quenching spectra for the samples of **P100**, **P150**, **P200**, and **P250**. Corresponding Stern–Volmer plots show non-linear behavior for **P35** and **P300** (Fig. 5b, e, and Fig. S7e–h (ESI†), respectively, for **P100**, **P150**, **P200**, and **P250**). Non-linear behavior is evident for both static and dynamic quenching of the perovskite.^{40,47} The concentration dilution effect was verified and is shown in Fig. S9 (ESI†).

To understand the charge transfer exciton behavior, time-correlated single-photon counting (TCSPC) studies are performed to observe the decay lifetime with the addition of graphite solution. In Fig. 6, the broad emission band peak at 512 nm shows a prolonged lifetime of nearly 25 ns. In general, the lifetime of charge-transfer excitons is observed by the electron–hole recombination rates directed by the Coulomb attraction force.

We should note that CTEs demonstrate various lifetimes when the electrons and holes are located on different energetic structures where the Coulombic attraction is changed by the local dielectric background. Emission decays of PNCs **P35** and **P300** at various ratios of PNC:graphite are presented in Fig. 6c and f, respectively. The decay could be fitted to the biexponential model for **P35** and triexponential for **P300**. The fitting

parameters are shown in Table 1. The average emission lifetime decreases with an increase in the ratio of PNC:graphite, which is in agreement with the steady-state quenching trend. The rapid electron injection process is responsible for a shorter lifetime than intrinsic excitons in semiconducting materials. We have observed a drastic change for the long component lifetime τ_2 , while there is no significant change in the short component τ_1 (Table 1) for sample **P35**. These observations suggest that charge transfer takes place predominantly in the presence of graphite solution. The following formula is used to calculate the average lifetime of the perovskite nanocrystal.

$$\tau_{\text{Avg}} = \frac{A_1\tau_1^2 + A_2\tau_2^2 + A_3\tau_3^2}{A_1\tau_1 + A_2\tau_2 + A_3\tau_3}$$

Table 1 TCSPC details of **P35** and **P300**

Graphite solution amount	A_1	A_2	A_3	τ_1	τ_2	τ_3	$\tau_{\text{Avg}}(\text{ns})$
P35							
0 mL	91.75	8.25	—	13.17	56.65	—	25
0.5 mL	85.75	14.25	—	11.03	36.44	—	20
1.0 mL	81.24	18.76	—	10.04	32.47	—	19
1.5 mL	84.30	15.70	—	10.68	33.78	—	19
2.0 mL	82.95	17.05	—	9.9	28.9	—	17
P300							
0 mL	927.34	7276.50	3601.30	24.60	6.43	1.49	12
0.5 mL	1091.90	7275.40	2828	22.04	6.11	1.48	11
1.0 mL	6342.30	4488.50	888.70	5.28	0.95	15.72	8
1.5 mL	424.73	7743.70	4109.20	19.42	4.46	1.07	7
2.0 mL	7633.30	411.53	4037.50	5.44	18.82	1.07	7

where A_1 , A_2 , and A_3 are the weight parameters, τ_1 , τ_2 , and τ_3 are the corresponding lifetimes and τ_{Avg} is the average lifetime.

The possible mechanism for the emission quenching of PNCs is energy transfer or electron transfer, or both. Here the energy transfer process may be ruled out because the absorption spectrum of graphite and the emission spectrum of PNC are negligible (Fig. S10 in ESI†). The only possible mechanism is the electron transfer from PNC to graphite. To investigate the dynamic behavior of charge transfer excitons, we have conducted pump-probe transient absorption spectroscopy on the pristine sample coupled with the nanocomposite. Herein, we have prepared two different compositions for transient absorption spectroscopy (TAS) investigations. Firstly, **P35** and the corresponding graphite composite **P35@G**, and the second one is **P300** and the corresponding graphite composite **P300@G**. Ultrafast transient absorption spectroscopy is based on the pump-probe technique where the experiment is progressed by excitation of molecules at a specific wavelength (called a pump) by a Ti:sapphire laser.

A weak pulse passing through the delay stage (called a probe) is used to investigate the dynamics of excited carriers. These ultrafast investigations are done using an oscillator (Micra), an amplifier (Coherent Legend), an optical parametric amplifier (OPA), and a spectroscopy system (Helios). The coherent legend has a fundamental beam of 800 nm (divided into two parts) with a repetition rate of 1 kHz, 35–45 fs pulse width, and average energy of 3–4 W. OPA generating tunable femtosecond (240–2600 nm) pulses with a pulse width of 70 fs

at the sample is used to stimulate the sample using the pump pulse. Surface Xplorer software and Origin 9.1 are utilized to accomplish fitting the data for analysis. The transient absorption behavior of MAPbBr₃ and after adding graphite was observed by modulating the pump laser in the excited state of the electron in the fs-ns time scale.

The photoinduced absorption, stimulation, and photoinduced bleaching are indicated in the spectral window of the TA spectra between 400 and 800 nm. Here, **P35** and **P35@G** were stimulated at 350 nm and 450 nm with an average pump power of 1 mW. The white light continuum probe pulse (delayed) *versus* pump pulse displays the absorption variations (ΔA) with wavelength and time. Fig. 7(a and b) illustrate the transient absorption spectra of **P35**. Here we get the ground state bleaching at ~ 480 nm, PIA at ~ 530 nm, and SE at ~ 550 to 800 nm probe delay wavelengths. The spectral characteristics after adding graphite demonstrate some modified signature, graphite added to **P35** quenches the bond energy and bond wavelengths as displayed in Fig. 7(c and d). The TAS characteristics indicate a wide spectral signal from 533 nm to 700 nm for the pristine compound (Fig. 7b). Interestingly this wide signal vanishes in the composite, suggesting the distinct dynamic behavior in the composite (Fig. 7d). A photobleaching signal is found with a growing component in the pristine sample demonstrating the ground state absorption characteristic. On the contrary, the growing component at the shorter time scale is eliminated with a quick decline in the composite clearly demonstrating the efficient charge transfer in the composite, as

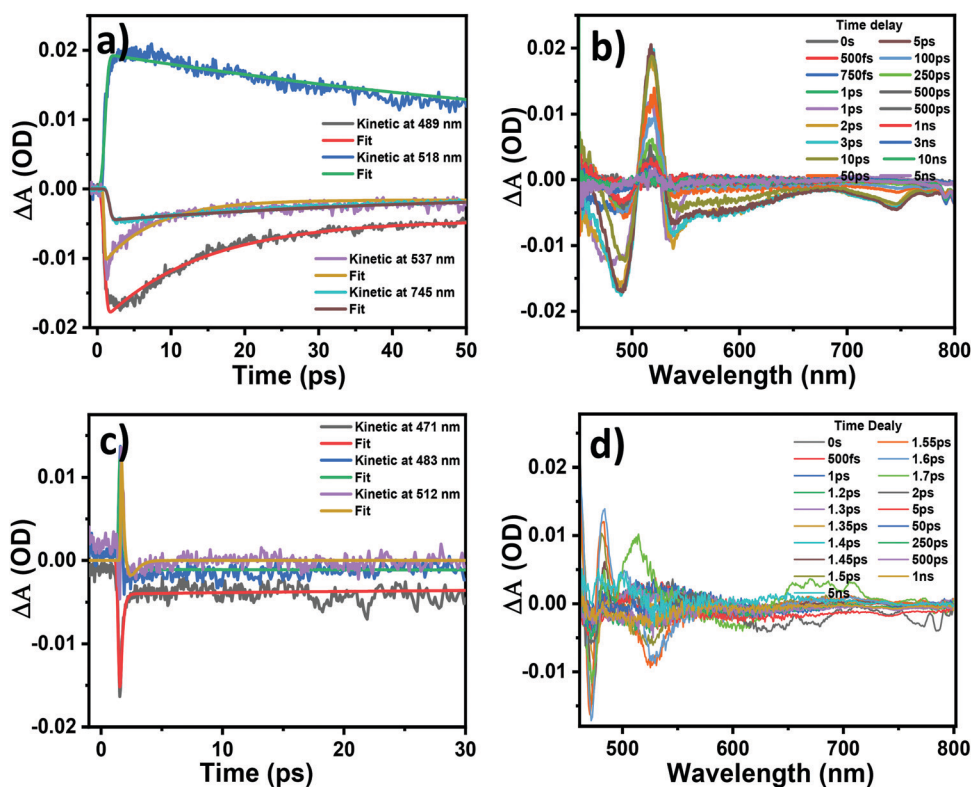


Fig. 7 The TAS spectra of (a) and (b) **P35**, and (c) and (d) **P35@G**.

illustrated in Fig. 7b and d. A significant signal from 503 nm to 533 nm in Fig. 7b is found, which entirely vanishes in the nanocomposite. By considering the fact that the excitons are evolved into the bleaching signals in the pristine sample, it may have a chance to contribute to the production of CTE as illustrated by the rapid decay kinetics in Fig. 7a and c. Therefore, the dynamics of excitons contain two processes. Firstly, an intrinsic exciton is created, which slowly fades with long decay kinetics. Secondly, the surface excitons readily transfer the charges to the nanocomposite showing effective creation of charge-transfer excitons at the interface. Similarly, we have done transient absorption spectroscopy investigations with the P300 sample and we found identical findings as for P35, as shown in Fig. S11 (ESI[†]). Details of TAS investigations for both the samples have been put in Table S3 (ESI[†]).

Conclusions

In summary, we have successfully synthesized perovskite nanocrystals of diverse shapes to examine the charge transfer potential with unfunctionalized graphite sheets. We have demonstrated that PNCs reasonably behave as an excellent electron donor when coupled with an acceptor such as graphite. Herein, we established effective charge transfer exciton creation, which efficiently transfers electrons from PNCs to graphite. TEM investigations demonstrate that self-aggregation happens with decreasing ligand concentration to generate different morphologies, while high ligand concentration gives small quantum dots. These nanocrystals have effectively bonded to the surface of the graphite sheets. This leads to the possibility that CTEs are formed in PNC graphite composites and provide opportunities to undertake sophisticated optoelectronics studies on the nanocomposite of PNC and graphite sheets.

Author contributions

Gaurav Kumar Nim: investigation, methodology, and writing – original draft. Sukanya Ghosh: writing – review and editing. Saurabh K Saini: transient absorption spectroscopy studies and editing. Mahesh Kumar: writing – review and editing. Prasenjit Kar: supervision, writing – review and editing.

Declaration

The authors declare that they have no known competing financial interests or personal relationships that could have influenced the work reported in this paper.

Conflicts of interest

The authors declare no conflict of interest.

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