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manipulation towards efficient thick organic solar
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Developing thickness-insensitive organic solar cells (OSCs) is of vital importance for meeting the requirements of the mass production of solar panels. Herein, the molecular aggregation behaviour of non-fullerene acceptors was manipulated via two different solvent additives, namely, 1-phenylnaphthalene (PN) and 1-chloronaphthalene (CN), to improve the charge transport in thick OSCs. The planar structure of the naphthalene ring coupled with large phenyl steric hindrance allows the PN additive to insert into the adjacent N3 molecules, disturbing their face-to-face π - π stacking and enhancing the J-aggregation in the blends. As a result, the PN-treated blend film exhibits red-shifted and broad absorption as well as increased charge carrier mobility, which is crucial for high-performance thick OSCs. In contrast, the CN additive, with a small side group possessing good solubility for N3 and a high boiling point, facilitates the H-type aggregation of acceptors in the spin-coating process. With the incorporation of 0.5 vol% PN, the 120-nm-thick D18-Cl:N3-based devices obtained a power conversion efficiency (PCE) over 18.0%, significantly higher than the 16.22% and 16.08% for the control and CN-treated devices, respectively. Impressively, when the active layer thickness is increased to 300 nm, the resultant device still obtained a PCE of 16.48%. This work provides a simple additive strategy to facilitate the precise control of molecular aggregation during the film formation process for high-performance thick OSCs.

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Introduction

Organic solar cells (OSCs) have garnered extensive attention in the realm of energy conversion technology because of their advantages of low-cost processing, flexibility and compatibility with large-area roll-to-roll manufacturing.^{1–4} A power conversion efficiency (PCE) of about 19% has been achieved in recent years, credited to advancements in novel donor and acceptor materials as well as device engineering optimization, signifying significant strides towards their commercial application.^{5–11} However, most efficient OSCs are typically fabricated with a thin active layer of 100–120 nm, which limits sufficient light harvesting and tends to form spot defects in scaled up manufacturing, leading to serious efficiency loss. In addition, the poor tolerance poses a critical challenge for roll-to-roll printing due to the inevitable thickness fluctuation during high-throughput manufacturing

processes.^{12–16} To address these issues and broaden the processing window for large-scale production, an active layer with thickness extended to several hundreds of nanometers could be more advantageous.^{16–19} Unfortunately, a thick active layer often leads to dramatic PCE reduction resulting from its unfavourable morphology and low mobility, causing a dilemma between the active layer thickness and device efficiency.^{20–23}

To improve the performance of thick active layer OSCs, considerable efforts have been devoted to material development as well as device engineering, such as the synthesis of temperature-dependent aggregative polymers,²⁴ the employment of layer-by-layer approach,¹⁷ and the introduction of fullerene liquid crystal into the active layer.²⁵ An effective strategy for thick-film regulation should ensure excellent crystallinity, optimal aggregation and superior vertical phase separation for efficient charge transport and collection.²⁶ Among these methods, the use of solvent additives within the active layer is deemed a convenient and effective approach for morphology regulation in thin-film OSCs.⁵ Nevertheless, the impact of solvent additives on the device performance of thick-film OSCs is still less understood. On the other hand, the molecular aggregation of donors and/or acceptors plays a pivotal role in film morphology and charge mobility in the active layers. In particular, non-fullerene acceptors (NFAs) usually form

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H- or J-aggregation, facilitating the building of a three-dimensional (3D) interpenetrated network structure that is conducive to realizing efficient charge transport in multiple directions.²⁷ Both molecular structures and processing conditions strongly affect the stacking behaviour during film formation. It is found that suitable additives can significantly promote the aggregation of non-fullerene small molecules, resulting in improved charge mobility in the bulk heterojunction (BHJ) films.²⁸ Although solvent additives have been widely used in highly efficient organic photovoltaics, their impacts on the aggregation kinetics are quite complicated and not clear enough. For a specific system, the selection of solvent additives is usually based on trial and error since the volatility, solubility and interaction with NFAs backbone or substituents potentially have a profound influence on the stacking styles of conjugated small molecules. It is still a challenge to develop an additive-choosing criterion for manipulating the molecular aggregation and regulating the film morphology.

Herein, we introduced two solvent additives with different steric hindrance, namely, 1-phenylnaphthalene (PN) and 1-chloronaphthalene (CN), to explore their influence on molecular aggregation and crystallization behaviour, charge transport and device performance in D18-Cl:N3-based OSCs. It is found that the addition of PN promotes J-aggregation since large steric hindrance from phenyl substituents disturbs the face-to-face π - π stacking of N3 molecules. In contrast, CN-treated films prefer to form H-type aggregation, which could be attributed to the small side group of CN. Although both PN and CN additives can improve the crystallization of N3 molecules due to the increased J- or H-type stacking, PN-treated films exhibit more favourable morphology with enhanced crystalline coherence length (CCL) in the D18-Cl:N3 blends. Compared to the D18-Cl:N3 devices with the CN additive, the

OSCs processed with PN additive show higher carrier mobility, higher charge collection efficiencies and reduced charge recombination. Meanwhile, the increased J-aggregation of N3 molecules in PN-treated D18-Cl:N3 blends is conducive to broaden the absorption spectra, resulting in an increased photocurrent. Benefitting from the enhanced charge transport and collection efficiency as well as the improved photon harvesting capability, a PCE of 18.0% is achieved for the PN-treated OSC devices with a 120 nm-thick D18-Cl:N3 active layer, significantly higher than that of the control (16.22%) and CN-treated (16.08%) devices, respectively. Moreover, the OSCs based on the 300 nm-thick PN-treated D18-Cl:N3 active layer still demonstrate an outstanding PCE of 16.48%, together with a short-circuit current density (J_{SC}) of 27.69 mA cm^{-2} , open-circuit voltage (V_{OC}) of 0.83 V, and fill factor (FF) of 71.98%. In contrast, the control and CN-treated devices with the same active layer thickness obtain moderate PCEs of 14.23% and 14.34%, respectively.

Results and discussion

The chemical structures of the polymer donor D18-Cl, non-fullerene acceptor N3, and solvent additives PN and CN are depicted in Fig. 1(a). With regards to the CN additive, PN has a large phenyl substituent that may induce distinct steric hindrance effects on the crystallization kinetics during the spin-coating process, thereby altering the molecular packing of donors or acceptors. To verify this hypothesis, we first investigated the optical properties of neat and blend films prepared with PN or CN additive. As shown in Fig. 1(b), the D18-Cl films processed without and with different additives exhibit similar UV-Vis absorption spectra from 450 to 600 nm, indicating the negligible impact of PN or CN on the polymer arrangement. However, the status is

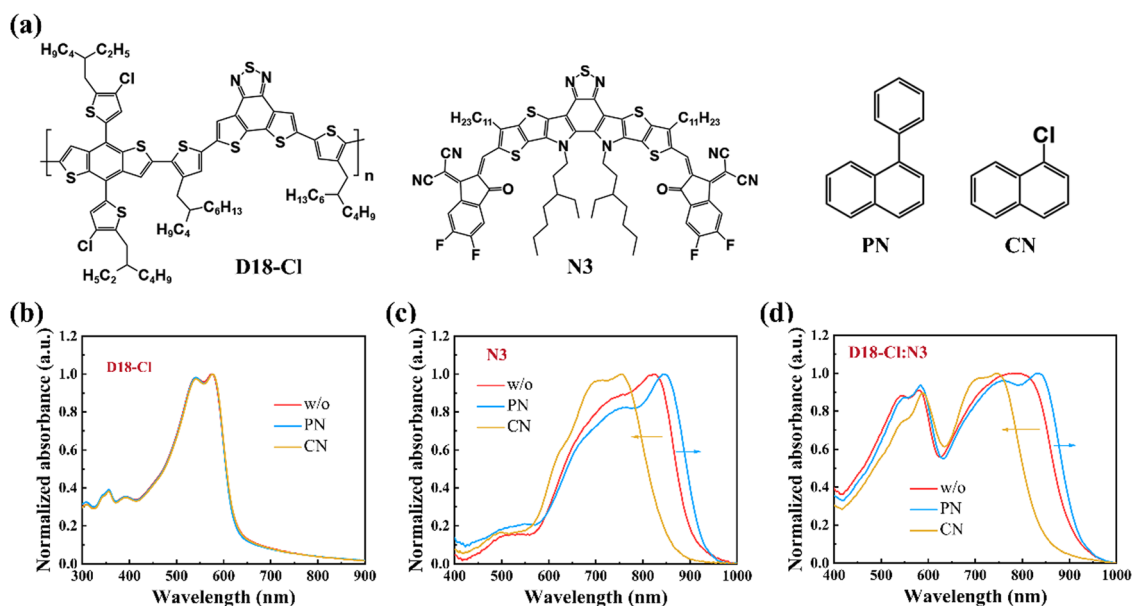


Fig. 1 (a) Chemical structures of D18-Cl, N3 and solvent additives PN and CN. (b)–(d) The absorbance of neat D18-Cl, neat N3 and D18-Cl:N3 blend films without and with different additives (0.5 vol%).



completely different for N3. As shown in Fig. 1(c), the pristine N3 film without any additives shows a primary absorption peak at 823 nm and a shoulder peak at 759 nm, which is assigned to the 0–0 and 0–1 vibronic transitions, respectively. The 0–0 peak of 0.5 vol% PN-treated N3 film undergoes a significant red-shift of about 22 nm and the 0–1 peak intensity (I_{0-1}) concurrently displays a proportional decrease, implying the enhanced J-aggregation of the N3 molecules. In contrast, the absorption of the 0.5 vol% CN-processed N3 film is dramatically blue-shifted by about 67 nm with a lower ratio of I_{0-0}/I_{0-1} , which suggests dominant H-aggregation in CN-treated N3 films. In addition, the PN-treated N3 films exhibit much broader absorption spectra with a full width at half maximum (FWHM) of 256 nm compared to the pristine and CN-treated N3 films, which is favourable for enhancing the photocurrent. This phenomenon was also observed in the PM6:Y6 blends processed with the solid additive and was attributed to the disturbed molecular packing of the acceptors caused by the solid additive.²⁸

Compared to the neat N3 films, the absorption spectra of the D18-Cl:N3 blends in Fig. 1(d) show the similar trend that PN-treated films exhibit an obvious red-shifted absorption while CN-treated ones are blue-shifted, which is attributed to the disparate aggregation behaviours of N3 in the blend films processed with different additives. Notably, D18-Cl:N3 blend films exhibit blue-shifted absorption compared to the corresponding pristine N3 films that could be ascribed to the restricted molecular motion of small molecules in the polymer donor matrix, thus contributing to the formation of H-type aggregation.²⁹ The influence of additive concentration on molecular stacking is shown in Fig. S1 (ESI†). With the increased concentrations of PN or CN, both neat N3 and D18-Cl:N3 blend films demonstrate gradually red- or blue-shifted absorption spectra, indicating a gradual increase of J- or H-aggregation in the films. The use of additives represents a strategy

to modulate the aggregation types, with the varied effects depending on the specific molecules involved.^{29–31} In N3-based systems, it can be concluded that J- and H-aggregation coexist while PN and CN are conducive to the dominant formation of J- and H-aggregation of N3, respectively. The different effects of PN and CN solvent additives on the molecular packing of N3 can be attributed to the distinct steric hindrance aroused by their substituents. A series of other naphthalene derivatives with various substituents were further employed as additives to confirm the impact of substituent hindrance effects on N3 molecular stacking. As shown in Fig. S2 (ESI†), the N3 films processed with 1,1'-binaphthalene additive show significantly red-shifted absorption peaks while the 1-methylnaphthalene, 1-methoxynaphthalene and 1-fluoronaphthalene-treated ones exhibit blue-shifted absorption spectra. Since 1-methylnaphthalene has similar polarity as PN, the blue-shifted absorption in the 1-methylnaphthalene-treated N3 film further confirms that it is the steric hindrance rather than the polarity of additives that significantly influences the aggregation behaviours of the N3 acceptors. Furthermore, Fourier transform infrared (FTIR) spectroscopy was employed to verify whether PN was removed, and the results are shown in Fig. S3 (ESI†). The additive PN exhibits two characteristic peaks at 3055 cm^{-1} and 1494 cm^{-1} corresponding to the C–H stretching vibration and aromatic ring skeleton deformation vibration, respectively. In the PN-treated N3 films, these characteristic peaks could be clearly observed, indicating that the additive still remains in the film after spin-coating due to its high boiling point. However, the characteristic peaks of PN disappear completely, implying that PN is effectively removed after spin-coating methanol treatment. Some other studies have also demonstrated that methanol treatment can remove the additives effectively.^{32–34}

Grazing incidence X-ray diffraction (GIXRD) measurements were performed to investigate the impact of additives on the

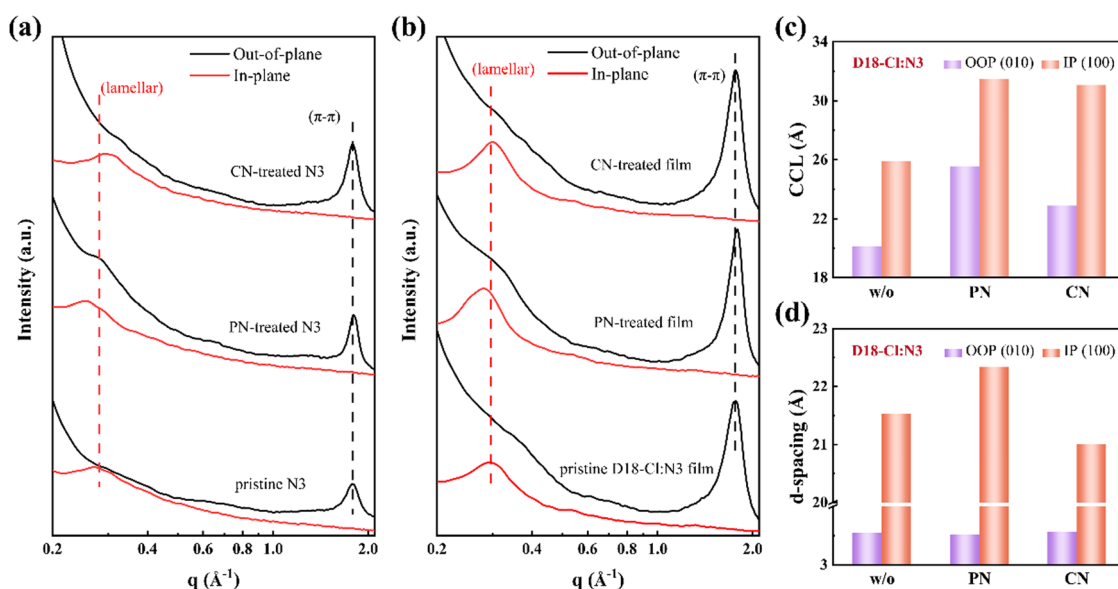
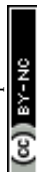


Fig. 2 (a) and (b) The GIXRD profiles of neat N3 and D18-Cl:N3 blend films without and with different additives (0.5 vol%). (c) The calculated CCLs and (d) d -spacing variation in the D18-Cl:N3 blend films without and with different additives (0.5 vol%).



molecular crystallinity in neat and blend films. The corresponding scattering profiles are compared in Fig. 2 and Fig. S4 (ESI[†]), and detailed quantitative analyses are summarized in Tables S1 and S2 (ESI[†]). The neat D18-Cl films without additives show a (010) π - π stacking peak along the out-of-plane (OOP) direction at $q_z = 1.679 \text{ \AA}^{-1}$ and a pronounced lamellar peak at $q_{xy} = 0.292 \text{ \AA}^{-1}$ along the in-plane (IP) direction in Fig. S4 (ESI[†]). The corresponding diffraction positions are unchanged for the D18-Cl films processed with PN or CN additives, indicating that both additives can hardly affect the polymer packing behaviours. In Fig. 2(a), the neat N3 film exhibits strong π - π stacking in the OOP direction and lamellar stacking in the IP direction due to their preferential face-on orientation with respect to the substrates. Compared to the pristine N3 film, the intensity of the π - π stacking peak for the PN- and CN-treated N3 films are significantly increased, implying better crystallinity and higher degree of molecular ordering. Meanwhile, the (010) diffraction peak at $q_z = 1.797 \text{ \AA}^{-1}$ in the PN-incorporated N3 films demonstrates closer π - π stacking among the N3 molecules. As depicted in Fig. 2(b), the D18-Cl:N3 blend film processed with 0.5 vol% PN shows a more

pronounced (100) lamellar diffraction peak at $q_{xy} = 0.281 \text{ \AA}^{-1}$ and (010) π - π stacking peak at $q_z = 1.787 \text{ \AA}^{-1}$. The crystalline correlation lengths (CCLs) of the D18-Cl:N3 films were calculated from the GIXRD profiles using Scherrer's equation,³⁵ and the results are plotted in Fig. 2(c). Compared to the blend film without the additive, the CCLs associated with lamellar stacking peaks in the CN- and PN-incorporated blend films were improved from 25.89 $\text{\AA}$ to 31.08 $\text{\AA}$ and 31.47 $\text{\AA}$, respectively, manifesting increased crystalline domains in the blend films. Simultaneously, the CCLs derived from the π - π stacking peak were boosted from 20.13 $\text{\AA}$ to 25.55 $\text{\AA}$, indicating an enhanced π - π ordering induced by the addition of PN. Additionally, the slightly decreased d -spacing in the OOP direction shown in Fig. 2(d) demonstrates the positive effect of PN in promoting the regular and compact π -stacking of N3 molecules in the blend films. The significant improvement of the molecular packing and crystallinity aroused by the PN additive is beneficial for charge transport in the resultant OSCs.

The atomic force microscopy (AFM) images of various films are depicted in Fig. S5–S8 (ESI[†]). The D18-Cl films show quite a smooth surface, and the root mean square (RMS) roughness

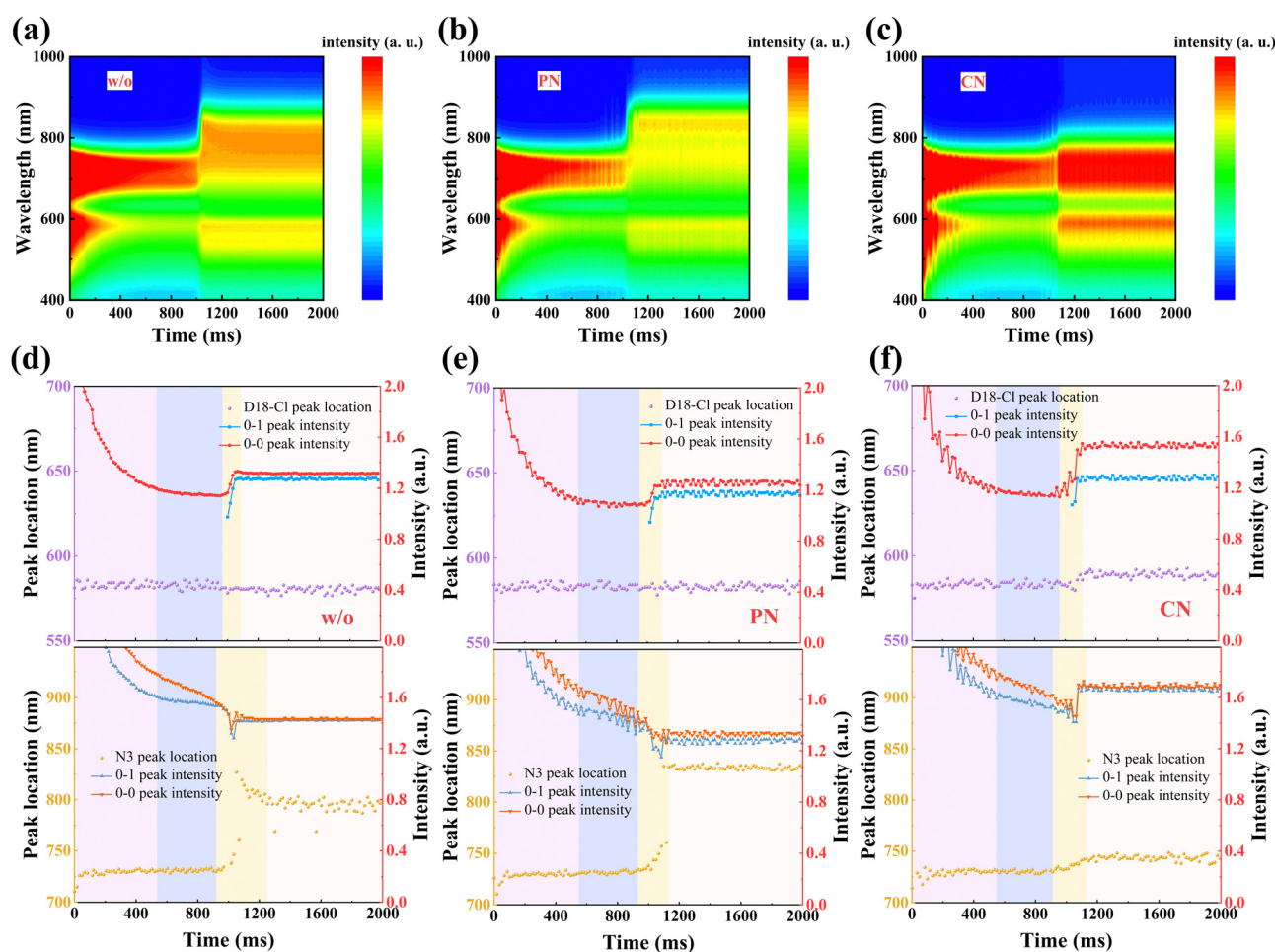
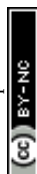


Fig. 3 Time-dependent contour maps of *in situ* UV-Vis absorption spectra of the D18-Cl:N3 samples processed (a) without any additives (b) with 0.5 vol% PN and (c) with 0.5 vol% CN. Time evolution of the peak location and intensity of the donor and acceptor in the D18-Cl:N3 samples processed (d) without any additives (e) with 0.5 vol% PN and (f) with 0.5 vol% CN.



marginally increased even with the volume fraction of PN or CN additives up to 2%. However, the neat N3 films processed with 2 vol% PN and CN additives show significantly increased RMS from 1.47 nm to 7.48 nm and 23.1 nm, respectively. These results validate again that the PN and CN additives primarily manipulate the aggregation of N3. The D18-Cl:N3 blend films processed with 0.5 vol% PN or CN form uniform nanofibril network structures, which are beneficial to exciton dissociation and charge transport. In detail, the RMS of 0.5% additive-processed films were slightly increased compared with the control film, which is attributed to the more ordered aggregation behaviour of the acceptors. When the CN additive is increased to 2 vol%, it will induce an abrupt expansion of phase separation, accompanied by the emergence of large-sized N3 aggregates. In comparison, a relatively small RMS of 2.57 nm is achieved in the D18-Cl:N3 film processed with 2 vol% PN, possibly implying the superior concentration tolerance of PN additives.

The formation process from the solution to the film state was further explored *via in situ* UV-Vis absorption spectra, as shown in Fig. 3(a)–(c), providing a comprehensive understanding of the drying and crystallization dynamics in D18-Cl:N3 blends. The characteristic absorption peak position and the intensity of the donors and acceptors were plotted as a function of spin-coating time in Fig. 3(d)–(f), respectively. According to previous reports,³⁶ the process can be divided into four temporal stages: (i) the fast removal of solution by centrifugal force, where only the absorption intensity drops quickly, but the peak position rarely changes; (ii) the evaporation of the main solvent persists while the additive remains in the “wet” film; (iii) the aggregation and crystallization of donors and/or acceptors, evidenced by either peak red shifting or absorbance improvement; (iv) film formation is eventually completed. During the whole spin-coating process, D18-Cl in all the three samples exhibits almost identical and constant peak positions due to similar pre-aggregation in solutions. The quickly increased peak intensity of D18-Cl at about 0.97–1.09 s mainly contributes to its high crystallization rate, which is not influenced by the additives. In the third stage of 0.92–1.25 s in Fig. 3(d), a non-monotonic transition is detected for the 0–0 peak position of N3 in the sample without additives, which is initially red-shifted and then blue-shifted until saturation at about 804 nm. This observation demonstrates that the J-aggregation of N3 molecules in the control sample obtained by fast solvent evaporation is kinetically *meta*-stable, which partly experiences a conversion to thermodynamically stable H-aggregation *via* self-organization. However, in the PN-treated samples shown in Fig. 3(e), the 0–0 peak position of N3 is rapidly red-shifted and kept at about 831 nm after 1.13 s, indicating suppressed molecular stacking transition. Meanwhile, the more balanced aggregation kinetics of D18-Cl and N3 in the PN-treated blend film predicts more balanced phase domains in the films, which is conducive to efficient charge generation and transport.^{17,37,38} Conversely, the CN-treated sample showed a nearly unchanged absorption peak and gradually increased 0–1 peak intensity at about 0.91 s after starting the spin-coating process. This can be

ascribed to the enhanced H-type stacking of the N3 molecules, possibly caused by the increased interactions between CN and N3, as discussed below.

In order to unveil the interactions between the additives and active materials, the Hansen solubility parameters (δ) of D18-Cl, N3, PN and CN were reviewed and are listed in Table S3 (ESI†).^{39–41} The solvent–molecular interaction parameters (χ) were calculated according to Flory–Huggins theory, and the results are as follows: $\chi_{\text{PN-D18-Cl}} = 1.04$, $\chi_{\text{PN-N3}} = 0.75$, $\chi_{\text{CN-D18-Cl}} = 0.51$, and $\chi_{\text{CN-N3}} = 0.42$. It is found that both PN and CN exhibit stronger interaction with the acceptor N3 than the donor D18-Cl, which is attributed to their π -conjugated structures that could enhance the van der Waals interactions with the end groups of NFAs.⁴² These interactions facilitate the insertion of additive molecules between adjacent acceptor tails and thereby create an additional space for relaxation and the subsequent ordered packing of NFA molecules. As a result, a reduced intermolecular π -stacking distance is obtained in the PN-treated D18-N3 films while the substantial steric hindrance associated with PN manipulates the dominant J-stacking arrangement. With regards to the CN additive, the lowest $\chi_{\text{CN-N3}}$ suggests its good solubility for N3, and the high boiling point of CN is conducive for the H-aggregation of N3 molecules because slower evaporation facilitates molecular reorganization into their thermodynamically stable state in the film formation process.

The molecular aggregation of NFAs are potentially effective in promoting charge transport in the active layers and thereby improving the device performance, especially for thick OSCs. Therefore, to examine the impact of different aggregation behaviours enabled by PN and CN on the photovoltaic performance, we fabricated thin and thick OSCs using the conventional architecture of ITO/PEDOT:PSS/D18-Cl:N3/PFN-Br/Al with an active layer thickness of 120 nm and 300 nm, respectively. The representative current density–voltage (J – V) characteristics under 100 mW cm^{−2} AM 1.5G illumination and EQE curves are shown in Fig. 4, and the corresponding device parameters are summarized in Table 1. In the case of OSCs based on a 120 nm-thick active layer without additive, a moderate power conversion efficiency (PCE) of 16.15% was achieved with a V_{OC} of 0.858 V, J_{SC} of 25.91 mA cm^{−2}, and FF of 73.26%. Upon the addition of 0.5 vol% PN, the PCE of the resultant device was enhanced to 18.02%, accompanied by a notably high J_{SC} of 27.68 mA cm^{−2}, V_{OC} of 0.855 V and FF of 76.42%. For comparison, the OSCs processed with 0.5 vol% CN only yield a maximum PCE of 16.08% with a V_{OC} of 0.860 V, J_{SC} of 24.86 mA cm^{−2} and FF of 75.24%. The calculated J_{SC} values from the EQE are in good agreement with the values achieved from the J – V measurements with a 5% mismatch. These results indicate that the J-aggregation of acceptors contributes to boosted J_{SC} while H-aggregation helps to enhance the V_{OC} values, consistent with the results reported.⁴³

In addition, the OSCs with varied volume fraction of PN from 0.25% to 2.0% show relatively high PCEs exceeding 16.02%, as shown in Fig. S9 (ESI†), indicating an outstanding additive concentration tolerance. More importantly, the PN-incorporated OSCs display superior thickness tolerance from 120 nm to 350 nm compared to the control and CN-treated



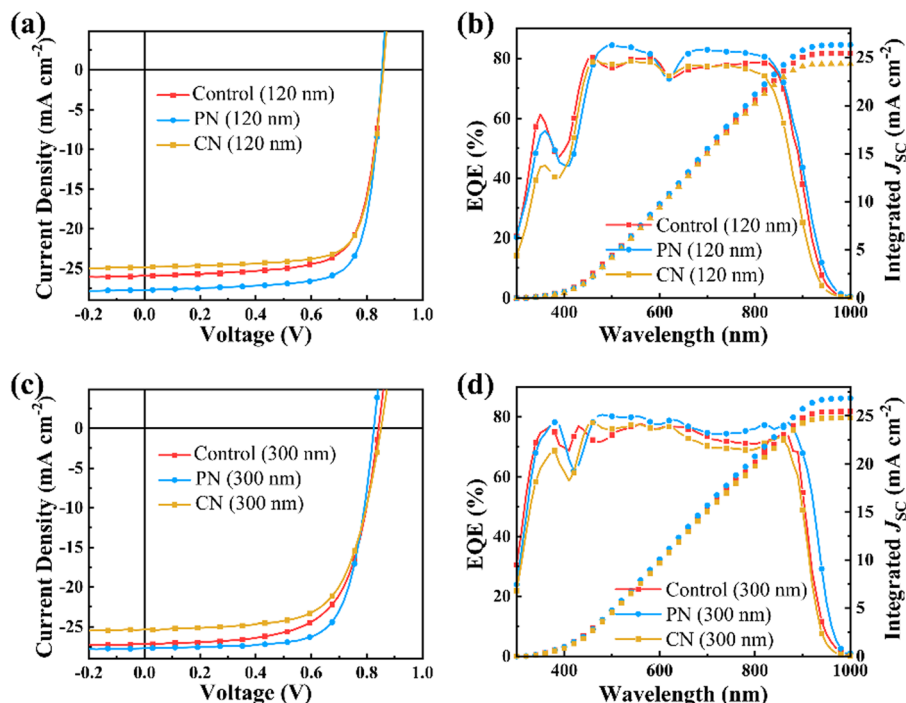


Fig. 4 (a) and (c) Illuminated J - V curves and (b) and (d) EQE spectra of the OSCs based on 120 nm or 300 nm-thick active layers with different additives (0.5 vol%).

Table 1 Photovoltaic parameters of the OSCs based on the 120 and 300 nm-thick active layers without and with different additives (0.5 vol%)

	Thickness (nm)	V_{OC} (V)	J_{SC} (mA cm^{-2})	J_{cal} (mA cm^{-2})	FF (%)	PCE ^a (%)
Control	120	0.858 (0.857 $\pm$ 0.006)	25.91 (25.76 $\pm$ 0.44)	25.44	72.98 (73.35 $\pm$ 1.01)	16.22 (16.19 $\pm$ 0.47)
PN	120	0.855 (0.844 $\pm$ 0.011)	27.68 (26.90 $\pm$ 0.88)	26.31	76.13 (75.87 $\pm$ 1.10)	18.02 (17.70 $\pm$ 0.35)
CN	120	0.860 (0.861 $\pm$ 0.003)	24.86 (24.66 $\pm$ 0.21)	24.32	75.24 (73.95 $\pm$ 1.37)	16.08 (15.58 $\pm$ 0.52)
Control	300	0.836 (0.839 $\pm$ 0.004)	26.46 (26.65 $\pm$ 0.49)	25.46	64.35 (63.98 $\pm$ 1.79)	14.23 (14.36 $\pm$ 0.68)
PN	300	0.827 (0.830 $\pm$ 0.010)	27.69 (27.77 $\pm$ 0.93)	26.49	71.98 (71.23 $\pm$ 0.75)	16.48 (16.38 $\pm$ 0.41)
CN	300	0.851 (0.846 $\pm$ 0.006)	25.32 (24.99 $\pm$ 0.57)	24.78	66.57 (66.07 $\pm$ 1.88)	14.34 (14.12 $\pm$ 0.71)

^a The average PCE in the brackets was calculated from 10 devices.

devices, as illustrated in Fig. S10 (ESI[†]). When the active layer thickness is increased to about 300 nm, the PCE of the PN-treated device is about 16.48%, along with a V_{OC} of 0.827 V, high J_{SC} of 27.69 mA cm^{-2} and FF of 71.98%, surpassing 14.23% and 14.34% for the control and CN-treated OSCs with the same active layer thickness, respectively. Both PN and CN additives can effectively improve the FF values while PN exhibits a more significant influence in thick OSCs, which is attributed to the more ordered molecular packing in the blend films. Moreover, the PN-treated device demonstrates remarkably increased J_{SC} , mainly due to the broadened absorption spectra aroused by the enhanced J-aggregation of the N3 molecules. Conversely, the V_{OC} of the PN-treated OSCs is decreased with regard to the CN-treated OSCs, which could probably be attributed to the dominant J-aggregation of N3 molecules.^{43,44} The synergistic improvement in the J_{SC} and FF values enables the outstanding performance of the PN-treated thick OSCs.

The charge transport and extraction were investigated to elucidate the superior performance of OSCs processed with PN,

especially for thick OSCs. The electron (μ_e) and hole mobility (μ_h) were measured *via* the space-charge limited current method, as presented in Table 2 and Fig. S11 in the ESI.[†] The μ_e and μ_h measured from the 120 nm PN-incorporated D18-Cl:N3 film were 1.67×10^{-3} and $1.35 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$, respectively, significantly outperforming the control and CN-incorporated ones. In the case of the 300-nm-thick active layers, the μ_h and μ_e were calculated to be 0.83×10^{-3} and $0.87 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ for the control blend, 1.21×10^{-3} and $1.36 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ for the PN-treated blend, and 1.07×10^{-3} and $1.00 \times 10^{-3} \text{ cm}^2 \text{ V}^{-1} \text{ s}^{-1}$ for the CN-treated blend, respectively. Obviously, both PN and CN additive could promote charge transport in the D18-Cl:N3 blend films regardless of their thickness due to improved molecular aggregation and crystallization. Compared to the CN additive, the PN-treated D18-Cl:N3 film demonstrates more pronounced enhancement of charge mobilities, contributing to higher J_{SC} and FF values in the resultant OSCs. The photocurrent density (J_{ph}) as a function of the effective voltage (V_{eff}) for these OSCs is plotted in Fig. 5.



Table 2 The charge mobilities of the active layers (120 nm and 300 nm) processed with different additives. The hole-only devices were fabricated with the structure of the ITO/PEDOT:PSS/active layer/MoO₃/Ag and the electron-only devices were prepared using the structure of ITO/ZnO/active layer/PFN-Br/Al

	Thickness (nm)	$\mu_h \times 10^{-3}$ (cm ² V ⁻¹ s ⁻¹)	$\mu_e \times 10^{-3}$ (cm ² V ⁻¹ s ⁻¹)	μ_h/μ_e
D18-Cl:N3	120	1.02	1.25	1.22
D18-Cl:N3-PN	120	1.35	1.67	1.24
D18-Cl:N3-CN	120	1.14	1.44	1.26
D18-Cl:N3	300	0.83	0.87	1.04
D18-Cl:N3-PN	300	1.21	1.36	1.12
D18-Cl:N3-CN	300	1.07	1.00	0.93

For all the devices, J_{ph} increases sharply with V_{eff} in the low effective voltage region and reaches saturation current (J_{sat}) at $V_{eff} \geq 3.0$ V. The exciton dissociation (η_{diss}) and charge collection efficiency (η_{coll}) given in Table S4 (ESI[†]) were estimated from J_{ph}/J_{sat} under short-circuit conditions and at the maximum power point, respectively.⁴⁵ It is noted that all the devices show nearly identical exciton dissociation efficiency, which is attributed to the suitable phase separation in the D18-Cl:N3 active layers. In the case of thin OSCs, the charge collection efficiency is slightly enhanced from 86.12% for the control device to 87.12% and 87.60% for the CN- and PN-incorporated devices, respectively. The PN-treated thick OSCs still show a η_{coll} of 87.30%, much higher than 80.56% and 84.79% for the control and CN-treated thick OSCs, respectively. These results indicate that the PN additive effectively enhances the charge transport and extraction behaviours, especially in the thick D18-Cl:N3 blend layer, thereby improving the photovoltaic performance of the corresponding thick OSCs.

The dependence of V_{OC} and J_{SC} on illuminated light intensity (P_{light}) was measured to estimate the charge recombination status in thin and thick OSCs. The relationship between J_{SC} and P_{light} can be described as $J_{SC} \propto (P_{light})^\alpha$, where α is the power-law exponent representing the extent of bimolecular recombination.⁴⁶ As displayed in Fig. 5, the extracted α values for the control, PN- and CN-incorporated thin OSCs were 0.96, 0.97, and 0.96, respectively. In case of the 300 nm-thick OSCs, the PN-treated device still exhibits the highest α value of 0.95 among the three kinds of devices. The dependence of V_{OC} on the natural logarithm of P_{light} gives a slope of nkT/q , where k is the Boltzmann constant, T is the temperature, and q is the elementary charge.⁴⁶ Alternatively, when monomolecular or trap-assisted recombination are involved, a strong dependence of V_{OC} on the light intensity with a slope close to $2kT/q$ will be observed. The OSCs with 120-nm-thick active layers exhibit similar slopes of 1.01–1.03 kT/q , indicating negligible trap-assisted recombination. However, when the active layer thickness is increased to 300 nm, the fitted slopes of the control and CN-incorporated devices are 1.15 and 1.16 kT/q , respectively, much higher than 1.09 kT/q for the CN-incorporated device. The suppressed trap-assisted and bimolecular recombination enabled by the PN additive contributes to a remarkable improvement in the photovoltaic performance of thick OSCs.

Transient absorption (TA) spectroscopy was employed to further assess the charge generation in D18-Cl:N3 blends

processed with different additives. The hole transfer process was firstly investigated by selectively exciting the acceptor N3 at a pump wavelength of 800 nm. The 2D color plots of the TA spectra for the neat films and D18-Cl:N3 blends pumped at 800 nm are shown in Fig. S12 and S17 (ESI[†]), and the TA spectra recorded at different time delays of the D18-Cl:N3 blends processed without and with the additives are depicted in Fig. S13 (ESI[†]) and Fig. 6(a), (b), respectively. Since D18-Cl cannot be pumped at 800 nm, the peak signal at 840 nm is ascribed to the ground state bleach (GSB) of N3 while the peak signal at 600 nm is assigned to the GSB of D18-Cl caused by the hole transfer from N3 to D18-Cl. Additionally, the excited-state absorption (ESA) signal for the local excited state of N3 is located at 930 nm, and a characteristic ESA signal after 1320 nm appeared upon optical excitation corresponding to an intermolecular delocalized excitation state due to the closely-packed N3 molecules. For the D18-Cl:N3 films, the fast decay of the N3 GSB signal after photoexcitation and a D18-Cl bleach signal that emerged within tens of picoseconds indicate the hole transfer process from N3 to the polymer D18-Cl.^{39,47} Fig. 6(c) shows the faster GSB increase of D18-Cl in the PN-treated D18-Cl:N3 blend, implying more efficient hole transfer between D18-Cl and N3. We retrieved the kinetic traces of three D18-Cl:N3 blends at 600 nm in Fig. 6(c) and Table S4 (ESI[†]) through bi-exponential function fitting, where τ_1 represents the ultrafast interfacial hole transfer process and τ_2 is the diffusion of excitons in the acceptor phase toward the interfaces.^{7,48} The PN-treated D18-Cl:N3 film shows a τ_1 of 0.54 ps and a τ_2 of 5.53 ps. It is noted that the τ_2 of the PN-treated D18-Cl:N3 film is shorter than 5.78 ps for the CN-incorporated blend film and 7.00 ps for the control blend film, manifesting faster exciton diffusion in the PN-treated films due to the improved molecular aggregation of N3. The almost same τ_1 values in the different D18-Cl:N3 blend films might be attributed to their dominant face-on orientation of the donors and acceptors, which promote rapid hole transfer at the D/A interfaces. The investigation then focused on the photophysical process occurring in various D18-Cl:N3 blend films under photoexcitation at 420 nm, at which the polymer D18-Cl is excited. As seen in Fig. S14 (ESI[†]), the neat D18-Cl film exhibits a GSB peak at ~ 600 nm and an ESA peak at 1200 nm instantly after photoexcitation. Due to the spectral overlap between D18-Cl photoluminescence and N3 absorption, as depicted in Fig. S16 (ESI[†]), it is predicted that a Förster resonance energy transfer (FRET) process may occur from the polymer D18-Cl to the acceptor N3. The 2D color plots of the TA spectra for the D18-Cl:N3 blends pumped at 420 nm are shown in Fig. S17 (ESI[†]), and the TA spectra recorded at different time delays are depicted in Fig. S15 (ESI[†]) and Fig. 6(d), (e), respectively. It is seen that the GSB of D18-Cl rapidly decays within the initial few hundred femtoseconds and then slowly increases within the ensuing timeframe of 1–50 ps. This phenomenon confirms that in addition to direct electron transfer, the FRET process also contributes to the charge generation in the D18-Cl:N3-based OSCs.^{49,50} Notably, the GSB signal of D18-Cl in the PN-treated D18-Cl:N3 films experiences a more pronounced resurgence, followed by a



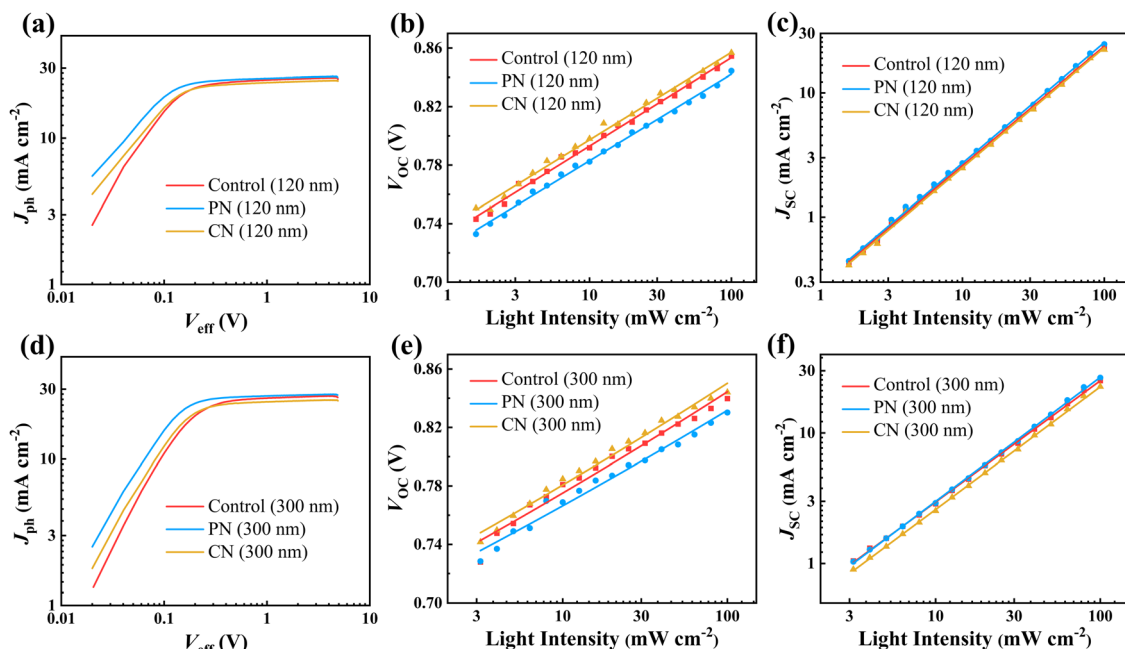


Fig. 5 (a)–(c) J_{ph} – V_{eff} curves and the dependence of V_{OC} and J_{SC} on the light intensity of the 120-nm-thick D18-Cl:N3 OSCs. (d)–(f) J_{ph} – V_{eff} curves and the dependence of V_{OC} and J_{SC} on the light intensity of the 300 nm-thick D18-Cl:N3 OSCs.

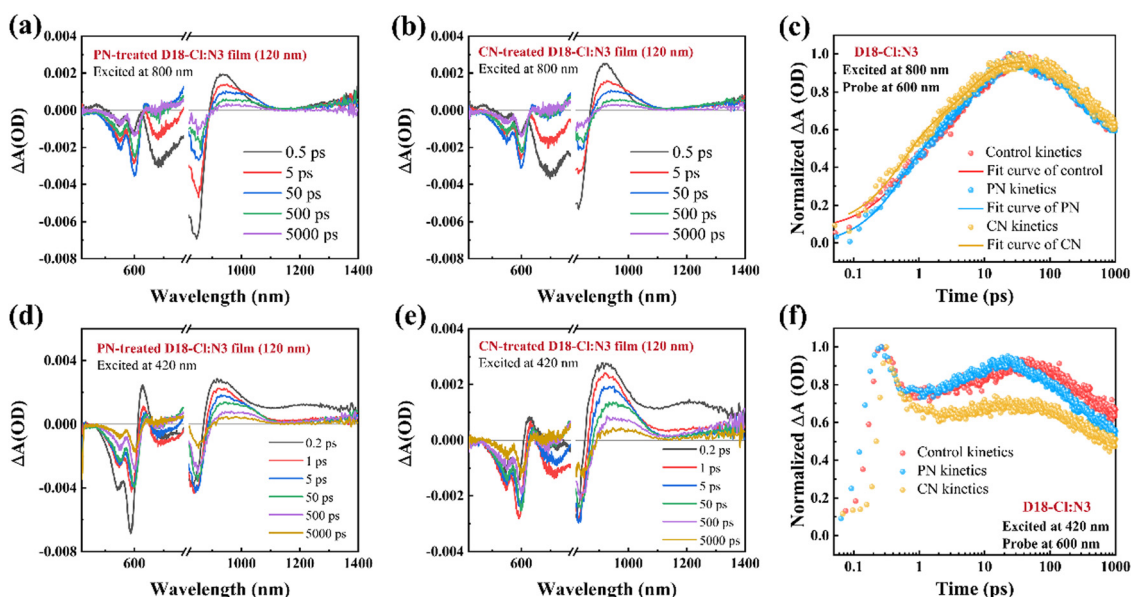


Fig. 6 (a) and (b) The line-cuts of fs-TA spectroscopy with pump at 800 nm at an indicated delay time of the PN- and CN-treated D18-Cl:N3 films. (c) The GSB kinetics of D18-Cl in different D18-Cl:N3 films excited at 800 nm. (d) and (e) The line-cuts of fs-TAS with pump at 420 nm at an indicated delay time of the PN- and CN-treated D18-Cl:N3 films. (f) The GSB kinetics of D18-Cl in different D18-Cl:N3 films excited at 420 nm.

faster decay after 100 ps, indicating that the PN additive accelerates the FRET process from the photoexcited D18-Cl to N3 as well as the subsequent reverse hole transfer process from N3 to D18-Cl. The efficient hole transfer and enhanced FRET process facilitate charge generation and result in superior photovoltaic performance in PN-treated D18-Cl:N3 OSCs.

Conclusion

In summary, we manipulated the molecular aggregation behaviour of non-fullerene acceptors *via* a facile additive strategy to enhance the photovoltaic performance of the OSCs based on thick active layers. Two additives of PN and CN with different substituent steric hindrances were used in the bulk



heterojunction D18-Cl:N3 active layer. It reveals that PN enables dominant J-aggregation of the acceptor N3 since the large phenyl substituents disturb the face-to-face π -stacking of adjacent N3 molecules. However, CN may enhance the H-aggregation of N3 acceptors during the film formation process due to its high boiling point and strong interaction with N3. Incorporation with PN or CN in the D18-Cl:N3 blend enhances molecular aggregations, leading to high charge mobilities that are crucial to thick OSCs. Moreover, the PN-treated D18-Cl:N3 active layer exhibits significantly red-shifted and broadened absorption spectra, which is beneficial for efficient light harvesting. The PN additive allows the OSCs based on 300-nm-thick D18-Cl:N3 active layer to achieve an efficiency of 16.48%, much higher than 14.23% of the corresponding devices without additive and 14.34% of the devices with the CN additive. This work provides an approach to regulate the molecular aggregation of non-fullerene acceptors towards high-performance thick OSCs.

Experimental

Materials

PEDOT:PSS (Clevios PVP AI 4083) was purchased from Heraeus Materials Technology Germany. D18-Cl and N3 used in this work were obtained from Yirou Photovoltaic Technology Co. Ltd. PFN-Br was obtained from Solarmer Material Inc. Chloroform (CF) was purchased from Xilong Scientific Co. Ltd. 1-Phenylnaphthalene (PN) was from Alfa Aesar Inc., and 1-chloronaphthalene (CN) was from TCI Shanghai. All the materials and solvents were used as received without further purification.

Device fabrication

The OSCs were fabricated with a conventional architecture of the ITO/PEDOT:PSS/active layer/PFN-Br/Al. The ITO glass substrates with a sheet resistance of $15 \Omega \text{ square}^{-1}$ were subjected to ultrasonic cleaning sequentially with detergent, deionized water, acetone, and isopropanol for 15 min of each step. The cleaned ITO glasses were then treated with UV-Ozone for 25 min before use. The hole transport layer of PEDOT:PSS ($\sim 25 \text{ nm}$) was spin-coated on ITO glass substrate and annealed at 140°C for 20 min in an oven. Then, the samples were moved into a glove box to spin-coat the active layer and the electron transport layer of PFN-Br. For the thin-film devices, the D18-Cl:N3 blend solution in CF (1:1.6 in weight) was stirred at 60°C for 30 min. CN or PN with a volume fraction of 0.5% was added and stirred for 30 min before spin-coating. In the preparation of the 300-nm-thick active layer films, a D18-Cl:N3 (1:1.6) solution with a concentration of 22.1 mg mL^{-1} was used. The solution was heated to 50°C and spin-coated at 1800 rpm. The 120-nm-thick active layer films were spin-coated from a D18-Cl:N3 (1:1.6) solution with the total concentration of 15.6 mg mL^{-1} at a higher rotational speed of 3000 rpm. The concentrations and spin-coating speeds used for other active layers are detailed below. For the 150 nm-thick active layer films, a concentration of 15.6 mg mL^{-1} was employed and the rotational speed was 2200 rpm. The 200-nm-thick active layer

films were prepared at a D18-Cl:N3 concentration of 18.2 mg mL^{-1} and a rotational speed of 2200 rpm. The 250-nm-thick films were produced with the same concentration of 18.2 mg mL^{-1} at a rotational speed of 1500 rpm. After spin-coating the D18-Cl:N3 film, the samples were thermally annealed at 90°C for 3 min. Subsequently, the PFN-Br solution in menthol (0.5 mg mL^{-1}) was spin-coated at 3000 rpm for 30 s. Finally, the Al cathode was thermally deposited with a shadow mask in a vacuum chamber. The active area of the cell was *ca.* 8 mm^2 .

Measurement and characterization

The UV-Vis absorption spectra were recorded using a UV-Vis Ocean Optics spectrophotometer. The AFM measurement was performed with a Bruker Dimension ICON AFM in the peak force mode. The thickness of the active layer was measured by a KLA Tencor P-7 surface profiler. TA spectra experiments were performed using a femtosecond pump-probe system. Illuminated current density-voltage (J - V) curves of the OSCs were recorded with a Keithley 2400 source meter under illumination of 100 mW cm^{-2} AM 1.5 G simulated solar light. The EQE of the devices was measured on an Enlitech QE-R spectral response measurement system.

Data availability

The authors declare that all data supporting the results reported in this study are available within the paper and the ESI.† Additional data used for the study are available from the corresponding author upon reasonable request.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

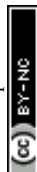
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References

- 1 Y. Wang, J. Yu, R. Zhang, J. Yuan, S. Hultmark, C. E. Johnson, N. P. Gallop, B. Siegmund, D. Qian, H. Zhang, Y. Zou, M. Kemerink, A. A. Bakulin, C. Müller, K. Vandewal, X. K. Chen and F. Gao, *Nat. Energy*, 2023, **8**, 978–988.
- 2 H. Liu, Y. Li, J. Wu, Y. Fu, H. Tang, X. Yi and Z. Xie, *J. Mater. Chem. C*, 2021, **9**, 9914–9921.



- 3 Z. Liu, Y. Fu, J. Wu, X. Yi, M. Zhao, M. Huang, J. Liu and Z. Xie, *Adv. Funct. Mater.*, 2024, **34**, 2401558.
- 4 P. Cheng, G. Li, X. Zhan and Y. Yang, *Nat. Photonics*, 2018, **12**, 131–142.
- 5 F. Sun, X. Zheng, T. Hu, J. Wu, M. Wan, Y. Xiao, T. Cong, Y. Li, B. Xiao, J. Shan, E. Wang, X. Wang and R. Yang, *Energy Environ. Sci.*, 2024, **17**, 1916–1930.
- 6 J. Fu, P. W. K. Fong, H. Liu, C. S. Huang, X. Lu, S. Lu, M. Abdelsamie, T. Kodalle, C. M. Sutter Fella, Y. Yang and G. Li, *Nat. Commun.*, 2023, **14**, 1760.
- 7 J. Wang, Y. Wang, P. Bi, Z. Chen, J. Qiao, J. Li, W. Wang, Z. Zheng, S. Zhang, X. Hao and J. Hou, *Adv. Mater.*, 2023, **35**, 2301583.
- 8 L. Kong, Z. Zhang, N. Zhao, Z. Cai, J. Zhang, M. Luo, X. Wang, M. Chen, W. Zhang, L. Zhang, Z. Wei and J. Chen, *Adv. Energy Mater.*, 2023, **13**, 2300763.
- 9 Z. Chen, H. Yao, J. Wang, J. Zhang, T. Zhang, Z. Li, J. Qiao, S. Xiu, X. Hao and J. Hou, *Energy Environ. Sci.*, 2023, **16**, 2637–2645.
- 10 X. Zheng, X. Wang, F. Sun, M. Wan, Z. Li, C. Xiao, T. Cong, T. Hu, Y. Liao and R. Yang, *Chem. Eng. J.*, 2023, **475**, 145520.
- 11 T. Xu, Z. Luo, R. Ma, Z. Chen, T. A. Dela Peña, H. Liu, Q. Wei, M. Li, C. E. Zhang, J. Wu, X. Lu, G. Li and C. Yang, *Angew. Chem., Int. Ed.*, 2023, **62**, e202304127.
- 12 C. Li, J. Zhou, J. Song, J. Xu, H. Zhang, X. Zhang, J. Guo, L. Zhu, D. Wei, G. Han, J. Min, Y. Zhang, Z. Xie, Y. Yi, H. Yan, F. Gao, F. Liu and Y. Sun, *Nat. Energy*, 2021, **6**, 605–613.
- 13 L. Zhu, M. Zhang, J. Xu, C. Li, J. Yan, G. Zhou, W. Zhong, T. Hao, J. Song, X. Xue, Z. Zhou, R. Zeng, H. Zhu, C. C. Chen, R. C. I. MacKenzie, Y. Zou, J. Nelson, Y. Zhang, Y. Sun and F. Liu, *Nat. Mater.*, 2022, **21**, 656–663.
- 14 K. Chong, X. Xu, H. Meng, J. Xue, L. Yu, W. Ma and Q. Peng, *Adv. Mater.*, 2022, **34**, 2109516.
- 15 F. Zhao, C. Wang and X. Zhan, *Adv. Energy Mater.*, 2018, **8**, 1703147.
- 16 Y. Chang, X. Zhu, K. Lu and Z. Wei, *J. Mater. Chem. A*, 2021, **9**, 3125–3150.
- 17 Y. Cai, Q. Li, G. Lu, H. S. Ryu, Y. Li, H. Jin, Z. Chen, Z. Tang, G. Lu, X. Hao, H. Y. Woo, C. Zhang and Y. Sun, *Nat. Commun.*, 2022, **13**, 2369.
- 18 T. Kirchartz, T. Agostinelli, M. Campoy Quiles, W. Gong and J. Nelson, *J. Phys. Chem. Lett.*, 2012, **3**, 3470–3475.
- 19 R. K. Gupta, R. Garai, M. Hossain, M. A. Afroz, D. Kalita and P. K. Iyer, *J. Mater. Chem. C*, 2021, **9**, 8746–8775.
- 20 H. R. Bai, Q. An, M. Jiang, H. S. Ryu, J. Yang, X. J. Zhou, H. F. Zhi, C. Yang, X. Li, H. Y. Woo and J. L. Wang, *Adv. Funct. Mater.*, 2022, **32**, 2200807.
- 21 L. Zhang, L. Hu, X. Wang, H. Mao, L. Zeng, L. Tan, X. Zhuang and Y. Chen, *Adv. Funct. Mater.*, 2022, **32**, 2202103.
- 22 Y. Zhu, X. Shen, H. Lai, M. Pu, Y. Zhu, X. Lai, S. Xiong and F. He, *Nano Energy*, 2023, **118**, 108991.
- 23 X. Song, K. Zhang, R. Guo, K. Sun, Z. Zhou, S. Huang, L. Huber, M. Reus, J. Zhou, M. Schwartzkopf, S. V. Roth, W. Liu, Y. Liu, W. Zhu and P. Müller Buschbaum, *Adv. Mater.*, 2022, **34**, 2200907.
- 24 H. Hu, P. C. Y. Chow, G. Zhang, T. Ma, J. Liu, G. Yang and H. Yan, *Acc. Chem. Res.*, 2017, **50**, 2519–2528.
- 25 F. Zhao, D. He, C. Zou, Y. Li, K. Wang, J. Zhang, S. Yang, Y. Tu, C. Wang and Y. Lin, *Adv. Mater.*, 2023, **35**, 2210463.
- 26 X. Song, H. Xu, X. Jiang, S. Gao, X. Zhou, S. Xu, J. Li, J. Yu, W. Liu, W. Zhu and P. Müller Buschbaum, *Energy Environ. Sci.*, 2023, **16**, 3441–3452.
- 27 H. Lai, Q. Zhao, Z. Chen, H. Chen, P. Chao, Y. Zhu, Y. Lang, N. Zhen, D. Mo, Y. Zhang and F. He, *Joule*, 2020, **4**, 688–700.
- 28 W. Zhang, Y. Wu, R. Ma, H. Fan, X. Li, H. Yang, C. Cui and Y. Li, *Angew. Chem., Int. Ed.*, 2023, **62**, e202309713.
- 29 Q. Zhao, X. Yu, Z. Xie, J. Liu and Y. Han, *ACS Appl. Energy Mater.*, 2018, **1**, 6338–6344.
- 30 Q. Zhao, J. Liu, H. Wang, M. Li, K. Zhou, H. Yang and Y. Han, *J. Mater. Chem. C*, 2015, **3**, 8183–8192.
- 31 J. Kim, K. J. Baeg, D. Khim, D. T. James, J. S. Kim, B. Lim, J. M. Yun, H. G. Jeong, P. S. K. Amegadze, Y. Y. Noh and D. Y. Kim, *Chem. Mater.*, 2013, **25**, 1572–1583.
- 32 L. Ye, Y. Jing, X. Guo, H. Sun, S. Zhang, M. Zhang, L. Huo and J. Hou, *J. Phys. Chem. C*, 2013, **117**, 14920–14928.
- 33 L. Xu, Y. Xiong, S. Li, W. Zhao, J. Zhang, C. Miao, Y. Zhang, T. Zhang, J. Wu, S. Zhang, Q. Peng, Z. Wang, L. Ye, J. Hou and J. Wang, *Adv. Funct. Mater.*, 2024, **34**, 2314178.
- 34 Y. Xie, H. S. Ryu, L. Han, Y. Cai, X. Duan, D. Wei, H. Y. Woo and Y. Sun, *Sci. China: Chem.*, 2021, **64**, 2161–2168.
- 35 S. Lilliu, T. Agostinelli, E. Pires, M. Hampton, J. Nelson and J. E. Macdonald, *Macromolecules*, 2011, **44**, 2725–2734.
- 36 L. Chen, R. Ma, J. Yi, T. A. Dela Peña, H. Li, Q. Wei, C. Yan, J. Wu, M. Li, P. Cheng, H. Yan, G. Zhang and G. Li, *Aggregate*, 2024, **5**, e455.
- 37 D. Luo, Z. Jiang, W. L. Tan, L. Zhang, L. Li, C. Shan, C. R. McNeill, P. Sonar, B. Xu and A. K. K. Kyaw, *Adv. Energy Mater.*, 2023, **13**, 2203402.
- 38 B. Lin, X. Zhou, H. Zhao, J. Yuan, K. Zhou, K. Chen, H. Wu, R. Guo, M. A. Scheel, A. Chumakov, S. V. Roth, Y. Mao, L. Wang, Z. Tang, P. Müller-Buschbaum and W. Ma, *Energy Environ. Sci.*, 2020, **13**, 2467–2479.
- 39 X. He, C. C. S. Chan, X. Zou, S. Zhang, P. W. K. Fong, J. Kim, G. Li, X. Hu, W. Ma, K. S. Wong and W. C. H. Choy, *Adv. Energy Mater.*, 2023, **13**, 2203697.
- 40 L. Ye, Y. Xiong, M. Zhang, X. Guo, H. Guan, Y. Zou and H. Ade, *Nano Energy*, 2020, **77**, 105310.
- 41 K. Yao, Y. X. Xu, X. Wang, F. Li and J. Yuan, *RSC Adv.*, 2015, **5**, 93689–93696.
- 42 L. Mei, X. Xia, R. Sun, Y. Pan, J. Min, X. Lu, A. K. Y. Jen and X. K. Chen, *Small*, 2024, **20**, 2305977.
- 43 Q. Zhao, H. Lai, H. Chen, H. Li and F. He, *J. Mater. Chem. A*, 2021, **9**, 1119–1126.
- 44 Q. Zhao and F. He, *J. Energy Chem.*, 2024, **93**, 174–192.
- 45 P. W. M. Blom, V. D. Mihailetschi, L. J. A. Koster and D. E. Markov, *Adv. Mater.*, 2007, **19**, 1551–1566.



- 46 S. He, Z. Shen, J. Yu, H. Guan, G. Lu, T. Xiao, S. Yang, Y. Zou and L. Bu, *Adv. Mater. Interfaces*, 2020, **7**, 2000577.
- 47 R. Wang, C. Zhang, Q. Li, Z. Zhang, X. Wang and M. Xiao, *J. Am. Chem. Soc.*, 2020, **142**, 12751–12759.
- 48 Z. Chen, T. Wang, Z. Wen, P. Lu, W. Qin, H. Yin and X. T. Hao, *ACS Energy Lett.*, 2021, **6**, 3203–3211.
- 49 F. D. Eisner, M. Azzouzi, Z. Fei, X. Hou, T. D. Anthopoulos, T. J. S. Dennis, M. Heeney and J. Nelson, *J. Am. Chem. Soc.*, 2019, **141**, 6362–6374.
- 50 Z. Chen, C. He, P. Ran, X. Chen, Y. Zhang, C. Zhang, R. Lai, Y. Yang, H. Chen and H. Zhu, *Energy Environ. Sci.*, 2023, **16**, 3373–3380.

