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Introduction

Dense metal–organic frameworks that adopt a perovskite-like architecture exhibit a wide range of fascinating properties¹ such as ferroelectricity,^{2,3} ferroelasticity,⁴ antiferromagnetic coupling^{5,6} and glassy phase transitions.^{7,8} In such ABX₃ materials, the (transition) metal B and the linker X form a ReO₃-like framework. The cation A, often a protonated amine, is then located within the open void of the cavity for charge balance (Fig. 1a). The hybrid perovskites [CH₃NH₃]PbI₃ and [NH₂CHNH₂]PbI₃,^{9,10} which started a new paradigm in the area of thin film solar cells, are intriguing examples.¹¹ Other examples include formate, azide and cyanide-based frameworks with the general formulas AB(HCOO)₃, AB(N₃)₃ and AB(CN)₃,

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Tuneable mechanical and dynamical properties in the ferroelectric perovskite solid solution $[NH_3NH_2]_{1-x}[NH_3OH]_xZn(HCOO)_3^{\dagger}$

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We report how mechanical and dynamical properties in formate-based perovskites can be manipulated by the preparation of an A-site solid-solution. In the series $[NH_3NH_2]_{1-x}[NH_3OH]_xZn(HCOO)_3$ with $x_{max} = 0.48$, the substitution of $[NH_3NH_2]^+$ by $[NH_3OH]^+$ is accompanied by a series of complex changes in crystal chemistry which are analysed using PXRD, SCXRD, ¹H solid state NMR, DSC and nanoindentation. NMR shows increased motion of $[NH_3NH_2]^+$ in $[NH_3NH_2]_{0.52}[NH_3OH]_{0.48}Zn(HCOO)_3$, which results in a shift of the ferroelectric-to-paraelectric phase transition temperature from $T_c = 352$ K (x = 0) to $T_c = 324$ K (x = 0.48). Additionally, the loss of hydrogen bonds directly influences the mechanical response of the framework; the elastic moduli and hardnesses decrease by around 25% from $E_{110} = 24.6$ GPa and $H_{110} =$ 1.25 GPa for x = 0, to $E_{110} = 19.0$ GPa and $H_{110} = 0.97$ GPa for x = 0.48. Our results give an in-depth insight into the crystal chemistry of ABX₃ formate perovskites and highlight the important role of hydrogen bonding and dynamics.

> respectively.¹²⁻¹⁴ In general, hybrid perovskites show strong structural similarities to their inorganic counterparts¹⁵ and well-established solid state principles were successfully applied to this emerging family.¹⁶ In particular, the large variation due to A, B or X site substitution permits the manipulation of bandgaps,¹⁷ dielectric responses,¹⁸ magnetic properties¹⁹ and so on. However, the chemical bond complexity in hybrid frameworks makes crystal engineering a challenging task, which, at the same time, can be seen as an opportunity to create frameworks with new and combined functionalities.

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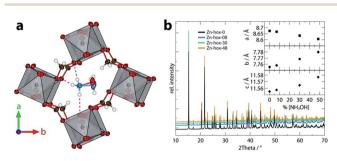


Fig. 1 View along the *c* axis of the pseudocubic unit cell of **Zn-hox-48** (a) with emphasis on the hydrogen bonds (purple) of the NH₃⁺-group and the different positions of the atoms O7 and N2 in the structure. In (b), the room temperature powder X-ray diffraction pattern of the compounds $[NH_3NH_2]_{1-x}[NH_3OH]_xZn(HCOO)_3$ with x = 0 (black), 0.08 (blue), 0.30 (green) and 0.48 (orange) are shown. The inset gives the evolution of lattice parameters along the solid solution. The lattice parameters were obtained from Pawley fits and fulfil Vegard's law. For Pawley fits and statistics, see ESI-Table 1 and ESI-Fig. 1.†

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Out of the many reported hybrid perovskites, formate-based materials show a notable large diversity and many different compounds with varying A-cations and metal species have been reported.^{12,20-22} Heterometallic hybrid formates are also known, *e.g.* [(CH₃)₂NH₂]Na_{0.5}Cr_{0.5}(HCOO)₃ and [C₂H₅NH₃]Na_{0.5}Fe_{0.5}-(HCOO)₃,^{23,24} which underline the compositional adaptability.

In all these compounds, the electronic configuration of the metal determines the low-temperature magnetic properties whereas the choice of the A cation, in particular the size, symmetry and molecular shape, is responsible for possible ferroelectric-to-paraelectric phase transitions. For instance, in $[NH_3NH_2]Zn(HCOO)_3$, the ferroelectric-to-paraelectric phase transition at 352 K is related to a temperature activated disorder of the NH₂-group between two positions. Interestingly, magnetoelectric coupling was observed in $[(CH_3)_2NH_2]Fe(HCOO)_3$, though the detailed nature of this phenomenon is still under debate.^{25,26}

Our work follows on from the important findings of Gao and co-workers who successfully prepared a solid solution between the two perovskite phases $[NH_3NH_2]Mn(HCOO)_3$ and $[CH_3NH_3]Mn(HCOO)_3$.¹⁸ Here, we describe the synthesis of the solid-solutions $[NH_3NH_2]_{1-x}[NH_3OH]_xZn(HCOO)_3$ with x = 0, 0.08, 0.30 and 0.48, where the end-members crystallise in a perov-skite-framework (x = 0, $[NH_3NH_2]Zn(HCOO)_3$) and a chiral channel structure (x = 1, $[NH_3OH]Zn(HCOO)_3$), respectively.²⁷⁻²⁹ Structural changes upon substitution are followed using powder and single crystal X-ray diffraction (PXRD and SCXRD) and differential scanning calorimetry (DSC), while ¹H magic angle spinning nuclear magnetic resonance spectroscopy (¹H MAS NMR) is used to access the dynamics of the A-cation. The changes in crystal chemistry and its impact on mechanical responses are then probed using nanoindentation.

Results and discussion

Synthesis and stoichiometry

Mild solution synthesis was used for the preparation of all compounds in this work. Starting from $NH_2NH_2 \cdot H_2O/NH_2$ -OH·H₂O or $NH_2NH_3Cl/HONH_3Cl$, HCOOH and $Zn(ClO_4)_2 \cdot 6H_2O$ in methanol, large crystals with sizes >0.2 mm are easily accessible (ESI†). Chemical analysis was then used to determine the chemical compositions of the products (Table 1). The lattice parameters obtained from PXRD data analysis (Fig. 1b) confirm that the solid-solutions obey Vegard's law. The lattice parameter

Table 1 Overview of $[NH_3NH_2]_{1-x}[NH_3OH]_xZn(HCOO)_3$ samples characterised in this work. The stoichiometry used during synthesis and real composition obtained from chemical analysis are given

| Name | $x \text{ in } [\text{NH}_3\text{NH}_2]_{1-x}[\text{NH}_3\text{OH}]_x\text{Zn}(\text{HCOO})_3$ | |
|-------------------------------------------------|------------------------------------------------------------------------------------------------|---------------------------|
| | (Synthesis) | (Chemical analysis) |
| Zn-hox-0 Zn-hox-08 Zn-hox-30 Zn-hox-48 | 0 0.30 0.70 0.80 | 0 0.08 0.30 0.48 |

a decreases along the series (-0.72%), whereas the lattice parameters *b* (+0.31%) and *c* (+0.19%) increase, leading to only a slight decrease of the volume from 777.30 Å³ to 775.16 Å³. The subtle change in lattice parameters is consistent with the similar sizes of $[NH_3NH_2]^+$, $r_{eff} = 217$ pm and $[NH_3OH]^+$ $r_{eff} = 216$ pm.¹⁶

Crystal chemistry and dynamics

In order to get further insights into the changes of crystal chemistry upon substitution, SCXRD of Zn-hox-48 ([NH₃- $NH_2]_{0.52}[NH_3OH]_{0.48}Zn(HCOO)_3)$ and Zn-hox-0 ([NH₃NH₂] Zn(HCOO)₃) at 120 K was performed. Similar to the parent compound **Zn-hox-0** (*Pna* 2_1 , a = 8.6706(2) Å, b = 7.72008(19) Å and c = 11.4872(3) Å), structure solution of **Zn-hox-48** was performed in the polar space-group *Pna2*₁ with lattice parameters, a = 8.61779(12) Å, b = 7.73073(10) Å and c = 11.50052(16) Å. The lattice parameters of Zn-hox-0 and Zn-hox-48 from SCXRD are consistent with the results from PXRD. In the electron density map, the oxygen atom of $[NH_3OH]^+$ could be clearly identified and was refined to an occupancy of 0.50, which is in good agreement with the results from chemical analysis. For the refinement, the total occupancy of the OH-group of [NH₃OH]⁺ and the NH₂-group of [NH₃NH₂]⁺ was set to 1, while the occupancies of both groups were refined freely. Detailed information about atomic positions and statistics of structure solution for Zn-hox-0 and Zn-hox-48 are given in ESI-Table 2-4.† In the final structural model for **Zn-hox-48**, the position of the $[NH_3NH_2]^+$ is effectively not changed in comparison to Zn-hox-0, leading to similar amine-cavity interactions.30 The oxygen atom of [NH₃OH]⁺ is slightly displaced with respect to the NH₂-group of $[NH_3NH_2]^+$. This displacement is related to the loss of one hydrogen bond (HB) and the subsequent maximisation of the remaining HB strength between the OH-group and the metalformate cavity; $d(O-H\cdots O) = 2.769(15)$ Å in comparison to the shortest N–H···O distance in **Zn-hox-0** d(N–H···O) = 2.998(2) Å. Additionally, a repulsion effect between the free lone pair at the oxygen atom and the negatively charged metal-formate cavity is possible and can support this displacement. Thus, the subtle shift of the [NH₃OH]⁺ axis with respect to the [NH₃NH₂]⁺ axis is a consequence of optimising the remaining amine-cavity interactions to increase the stability of the structure. Since the end members of the phase diagram crystallise in two different solution structures, along the solid crystal going [NH₃NH₂]_{1-x}[NH₃OH]_xZn(HCOO)₃ there is no driving force present that represents a transition from one distorted perovskite architecture to another. This is consistent with the small changes of trans Zn–Zn–Zn angles along c of 170.04°/169.79° and average Zn–O distances of d(Zn–O) = 2.103/2.104 Å for Zn– hox-0/Zn-hox-48, respectively.

The impact of substitution on the ferroelectric-to-paraelectric phase transition was analysed using DSC. With increasing *x*, the phase transition temperature (T_c) decreases linearly from $T_c = 352$ K for **Zn-hox-0** to $T_c = 324$ K for **Zn-hox-48** (Fig. 2). For **Zn-hox-0**, the motion of the NH₂-group is hindered by HBs formed between the NH₂-group and the metal-formate cavity. Additionally, the free electron pair located at the NH₂-

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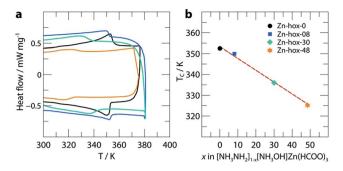


Fig. 2 DSC data (a) and phase transition temperatures (b) as a function of x in $[NH_3NH_2]_{1-x}[NH_3OH]_xZn(HCOO)_3$. The dotted red line acts as a guide for the eye.

group points into free space to reduce electrostatic repulsion effects with the negatively charge metal-formate framework. The subtle distortion of the metal-cavity described above seems to perturb this balance and a motion of the NH₂-group becomes more favourable than localised HBs. This is supported by the DSC signal, which flattens out and becomes thermally dispersed with increasing *x*. A similar trend was found for the decomposition temperature ($T_{\rm D}$), which decreases from $T_{\rm D}$ = 394.9 K to $T_{\rm D}$ = 392.3 K for **Zn-hox-0** and **Zn-hox-48**, respectively (ESI-Fig. 2†).

To investigate the impact of $[NH_3NH_2]^+$ substitution on the dynamics and T_c in more detail, **Zn-hox-48** was analysed using ¹H solid state NMR. The spectrum shows similar resonances to that of **Zn-hox-0** arising from the HCOO⁻ ($\delta^{1}H = 8.4$ ppm) and NH₃NH₂ groups ($\delta^{1}H = 9.3$ ppm for NH₃ and 4.0 ppm for NH₂),³⁰ while an additional resonance is observed at $\delta^{1}H = 10.4$ ppm (Fig. 3a). The assignments of the NH₃OH groups were made on the basis of a 2D double-quantum NMR experiment (Fig. 3b and ESI-Fig. 3 and 4† for details). The measured peak intensity ratios of 2.6 : 4.1 : 6.0 : 2.1, (NH₃OH : NH₃NH₂ & NH₃OH : HCOO : NH₃NH₂) are in good agreement with the

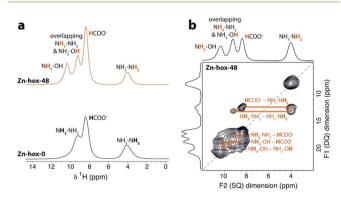


Fig. 3 (a) Experimental ¹H NMR (11.7 T) spectra of Zn-hox-0 and Zn-hox-48 are shown with an MAS frequency of 50 kHz. The chemical shifts and integrated intensities for Zn-hox-48 are given in the text. At 50 kHz MAS the average sample temperature is ~307 K, as determined from calibration experiments on lead nitrate.³¹ (b) ¹H 2D doublequantum spectrum (11.7 T) of Zn-hox-48 at a MAS rate of 60 kHz. Orange lines show the different correlations and indicate hydrogen atoms that are close in space.

expected ones for x = 0.48 (2.9 : 4.1 : 6.0 : 2.1), confirming the stoichiometry obtained by chemical analysis. Closer inspection of the NMR spectra in Fig. 3a shows that the peak linewidths are narrower for the Zn-hox-48 compound than the Zn-hox-0 compound. For example, the full-linewidth at half-maximum intensity (FWHM) of the hydrazinium -NH2 peak decreases from 1.07 ppm in Zn-hox-0 to 0.80 ppm in Zn-hox-48. Experiments at different magnetic field strengths and different MAS speeds (see ESI-Fig. 4[†]) confirm that the peak linewidths are dominated by dipolar interactions between the spins. In this regime, motion of the hydrogen containing groups averages the ¹H dipolar interactions, with faster motion giving rise to narrower NMR peaks.³² We can therefore conclude that the hydrazinium cations are more mobile in Zn-hox-48 than in Znhox-0. This observation is consistent with results from DSC analysis and the above mentioned decrease of $T_{\rm c}$. Hence, the introduction of hydroxylammonium cations in the ReO3-like cavities has an influence on the dynamics of hydrazinium cations located in other cavities. This finding points to the presence of cooperative effects mediated through the 3D metalformate framework, even though X-ray diffraction revealed only minor changes in the metal-formate cavity upon substitution. A similar mechanism, where the metal-formate cavity plays a key role, is currently discussed for magnetoelectric and magnetoelastic coupling pathways.²⁵

Mechanical properties

Nanoindentation was used to study the mechanical properties for all compounds in this work, see ESI[†] for experimental details. Mechanical responses were measured at ambient temperatures along [110], which corresponds to the [100] direction of the pseudocubic perovskite unit cell (Fig. 1a). Along the solid solution, a significant decrease of the elastic moduli from $E_{110} = 24.6$ GPa for **Zn-hox-0** to $E_{110} = 19.0$ GPa for **Zn-hox-48** is observed (Fig. 4a). Similarly, the hardness decreases from $H_{110} = 1.25$ GPa for **Zn-hox-0** to $H_{110} = 0.97$ GPa for **Zn-hox-48** (Fig. 4b). For more detailed results from nanoindentation, see ESI-Table 5[†] and ref. 30. The change of approx. 25% for *E* and *H* can be linked to the chemical manipulation of amine–cavity

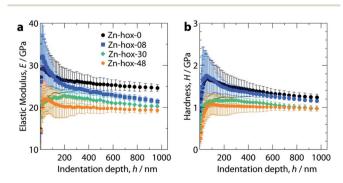


Fig. 4 Elastic moduli, *E* (a) and hardness, *H* (b) as a function of indentation depth, *h* along [110]. With increasing *x* in $[NH_3NH_2]_{1-x}$ - $[NH_3OH]_xZn(HCOO)_3$, the elastic modulus and hardness decrease monotonically due to reduced hydrogen bonding interactions upon substitution.

interactions across the series. The dependency of *E* and *H* with *x* seems to be linear to a good approximation, which is consistent with PXRD, DSC and TGA data (ESI-Fig. 4 \dagger).

In a previous study we have shown that amine-cavity interactions in Zn-hox-0 are relatively strong and we identified salt bridge-like interactions between the metal-formate cavity and the NH₃-group, as well as two hydrogen bonds, medium in strength, formed by the NH₂-group of [NH₃NH₂]⁺.³⁰ By gradually replacing $[NH_3NH_2]^+$ by $[NH_3OH]^+$, the number of HBs from the A cation to the metal-formate cavity is effectively reduced. This can be seen as a first order effect which is responsible for a relatively large decrease in the mechanical stability of the framework. The displacement of the [NH₃OH]⁺ discussed above, together with the more polarized O-H bond, strengthen the newly formed HB between the OH-group and the cavity. This partially compensates for the weakening of the framework caused by the first order effect. Additionally, the more polar N-O bond in $[NH_3OH]^+$ and the subsequent strengthening of amine-cavity interactions of the NH₃-group can play a role. Thus, the change in mechanical properties can be linked to the replacement of [NH₃NH₂]⁺ with [NH₃OH]⁺, which has a complex influence on the crystal chemistry of the compound.

Conclusion

In conclusion, we have shown how the dynamics, phase transition temperatures and mechanical properties can be tuned in formate perovskites by using the solid-solution approach. Within the series $[NH_3NH_2]_{1-x}[NH_3OH]_x Zn(HCOO)_3$ (x = 0, 0.08, 0.30 and 0.48), the progressive substitution of $[NH_3NH_2]^+$ by $[NH_3OH]^+$ induces a subtle but complex change in crystal chemistry. The variation of one structural parameter, i.e. the removal of HBs, induces a cascade of small distortions which then significantly influence the properties of the compounds. ¹H solid state NMR indicates higher mobility of [NH₃NH₂]⁺ in Zn-hox-48 with respect to Zn-hox-0. Consequently, the ferroelectric-to-paraelectric phase transition temperature decreases from $T_{\rm c} = 352$ K for **Zn-hox-0** to $T_{\rm c} = 324$ K for **Zn-hox-48**, because the threshold temperature where the entropy gain overcomes the loss in enthalpy decreases. A more direct effect of the loss of HBs upon substitution is the increased compliance of the framework. This is reflected in the elastic moduli and hardnesses of the systems which decrease from $E_{110} = 24.6$ GPa and $H_{110} = 1.25$ GPa for **Zn-hox-0** to $E_{110} =$ 19.0 Gpa and $H_{110} = 0.97$ GPa for **Zn-hox-48**. This confirms previous results in which the sensitivity of mechanical properties towards amine-cavity interactions was demonstrated.33 The variation of A-cation dynamics, however, is complex and underlines the challenges materials scientists face when preparing new materials with targeted functionalities. The results presented here are consistent with the solid solution $[NH_3NH_2]_{1-x}[CH_3-$ NH₃]_xMn(HCOO)₃ reported by Gao and co-workers.¹⁸ We have gone beyond their work by studying the A cation dynamics by ¹H NMR and the mechanical properties using nanoindentation. Additionally, by using a protonated amine as a substituent that does not form a perovskite structure in its pure form, our work opens a new dimension in tuning properties of formate perovskite frameworks.

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