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Exceptionally flexible quinodimethanes with multiple conformations: polymorph-dependent colour tone and emission of crystals

Extremely flexible quinodimethanes! Since polyaryl-substituted quinodimethanes can adopt multiple conformations such as folded, twisted, planar, and twisted-folded forms, a drastic change in the colour tone and emission of the crystals was observed depending on the electronic configurations of the conformers.

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## Exceptionally flexible quinodimethanes with multiple conformations: polymorph-dependent colour tone and emission of crystals†

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Tetraarylated quinodimethanes, which are one of the overcrowded ethylenes, can adopt multiple conformations, such as folded, twisted, and other forms. However, it is difficult to obtain all of these conformations due to energy differences among them. To tackle this issue, we designed flexible polyaryl-substituted tetraazaanthraquinodimethanes (N<sub>4</sub>AQDs), in which a steric hindrance in the overcrowded fjord region was moderately reduced. As a result, X-ray analyses revealed that N<sub>4</sub>AQDs could adopt not only folded and twisted forms but also intermediate structures, e.g., planar and twisted-folded forms, in pseudopolymorphs. Since there is a large difference in the electronic configuration among conformers, the colour tone of crystals varied from yellow to red depending on the conformation of overcrowded N<sub>4</sub>AQDs. Notably, red to near-infrared emission was observed for the crystal containing twisted conformers. This is the first example of exceptionally flexible quinodimethanes with tunable photophysical properties and multi-chromic behaviour based on multiple conformations.

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## Introduction

Stereoisomers (e.g., enantiomers and diastereomers) with topologically equivalent structures but different conformations and configurations often exhibit inherent physical properties. Since some diastereomers can be switched among two or more stable states by macroscopic external stimulation such as by light and/or heat, the microscopic properties based on the conformation/configuration can be controlled.<sup>1–3</sup> These have been widely studied for potential applications in molecular switches, memory devices, and sensor materials.<sup>4–9</sup>

With regard to such response systems, the conformations of overcrowded ethylenes (OCEs), in which a C=C double bond is surrounded by bulky substituents, can be switched between two or more conformers [e.g., folded (F) and twisted (T) forms] by external stimuli,<sup>10–13</sup> especially for use in molecular motors.<sup>14–18</sup> Generally, the T-form has a higher HOMO and lower LUMO than the F-form, resulting in a significant difference in absorption and emission wavelengths between conformers. Recently, there

has been steady progress toward obtaining the T-form in addition to the F-form for their potential application as stimuli-responsive molecules;<sup>19–26</sup> however, investigations on the T-form remain insufficient due to its potential instability or shorter lifetime.

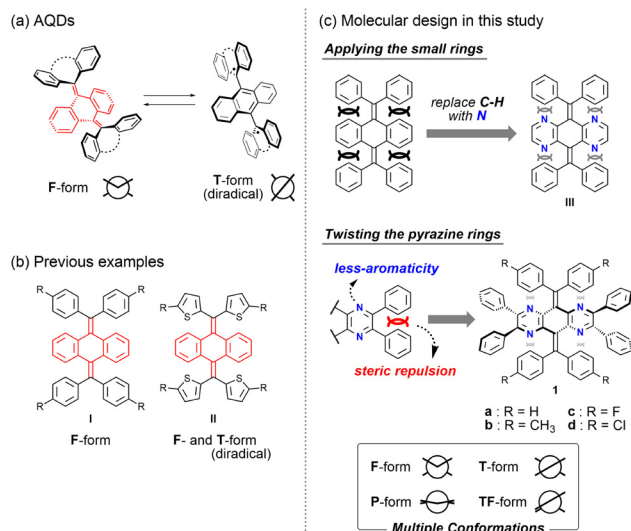
Among various OCEs, anthraquinodimethanes (AQDs) have thus far been used as a common way to obtain the T-form in addition to the F-form by introducing tricyclic aromatic units (Fig. 1a). Notably, since the T-form as seen in AQDs is the open-shell diradical species, various properties such as colour and magnetic properties can be controlled by external stimuli based on a change in the conformation between the closed-shell F-form and the open-shell T-form.<sup>27,28</sup> We also reported several tetraaryl-AQD derivatives such as **I**<sup>29</sup> and **II**<sup>12</sup> (Fig. 1b). In particular, for **II**, we have realised the control of redox properties based on the thermal equilibrium between the closed-shell F-form and the open-shell T-form. As shown in these AQDs, while open-shell species are interesting from the viewpoint of their potential applications, their high reactivity often causes a problem in handling. Therefore, we envisaged that further functional molecules that exhibit drastic changes in absorption and emission based on the conformers with different electronic configurations should be obtained when quinodimethane derivatives could exhibit the “closed-shell” T-form in addition to the F-form.

Since it is difficult to isolate more than two structures for the same molecule in OCEs, there are very few examples where the existence of both F- and T-forms in (pseudo)polymorphs was

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**Fig. 1** (a) Structural change of AQDs. (b) Previous examples of AQDs. (c) Molecular design in this study to make multiple conformations stable by decreasing the steric hindrance between aromatic rings (smaller aromaticity).

confirmed by X-ray analyses.<sup>30–33</sup> As exemplified by these molecules, dramatic chromic behaviour based on the difference in the electronic configuration between F- and T-forms is expected; however, the molecular design strategy is still limited to such molecules with rigid and planar aromatic groups. Since these molecules are very attractive, we envisaged that we could obtain unexplored conformations in addition to these two conformers by an appropriate molecular design. Therefore, we aimed to construct a unique molecule that could adopt not only F- and T-forms but also other conformations [e.g., planar (P) and twisted-folded (TF) forms], so that unprecedented switching behaviour could be achieved with the involvement of multiple states.

Herein, we report tetraazaanthraquinodimethane ( $N_4$ AQD) derivatives **1a–d**, which can adopt both the closed-shell T- and F-forms in crystals by reducing the steric hindrance in the overcrowded fjord region. Furthermore, we have also succeeded in creating not only these two structures, but also several intermediate structures in polymorphs. In other words, we modulated the photophysical properties based on the conformation observed in the crystal. Generally, since dynamically averaged physical properties are observed in solution, the present results represent the first example where colour tone and emission colour are controlled based on multiple conformations in the crystal that can be generated by recrystallization from appropriate solvents.

## Results and discussion

### Design strategy

On the basis of the difference in structures found in quinodimethane derivatives **I** and **II**, we expected that the conformations could be significantly affected by a change in steric

hindrance around the central C=C double bond. Therefore, we considered that the decrease in steric hindrance in the fjord region would be key to creating the T-form for AQDs. At the same time, to control the conformations of AQDs, we propose a new strategy: reducing the steric hindrance in the overcrowded fjord region by replacing the C-H moiety of the central AQD skeleton with nitrogen atoms (derivative **III**, Fig. 1c). Instead of an approach that uses thienyl groups, which contribute to the stabilization of open-shell species, the introduction of pyrazine rings into the central AQD skeleton is designed to reduce the aromaticity of the tricyclic core and allow the skeleton itself to twist.

Before starting the synthesis, we performed density functional theory (DFT) calculations at the (U)B3LYP/6-31G\* level for  $N_4$ AQDs **III** (as well as **I** and **II** for comparison). The results showed that the closed-shell T-form in **III** is still calculated as a transition state with one imaginary frequency as well as **I** (Table 1). Thus, to obtain the T-form as a stable structure, we decided to use an additional approach to perturb the molecular structure by the introduction of two *vic*-diphenyl groups into both sides of the central  $N_4$ AQD skeleton (compound **1**, Fig. 1c). This strategy can induce a further twisting of the less-aromatic pyrazine ring due to steric repulsion between adjacent phenyl groups, resulting in a decrease in steric hindrance between the  $N_4$ AQD core and the aryl groups on the exocyclic bonds to make the T-form more stable. DFT calculations were then performed on the  $N_4$ AQD derivative **1**, which suggested that the T-form would be obtained as the most stable structure and the F-form would also exist as a metastable structure (Table 1 and Table S9, ESI†). The results of the DFT study for non-aza analogue **IV** confirmed the validity of our molecular design. Thus, we focused on  $N_4$ AQDs with peripheral phenyl groups, with the expectation that both the F- and T-forms could appear in crystals and designed four derivatives **1a–d** to investigate the effect of the steric and electronic properties of the substituent at the 4-position on each of the aryl groups (Table S6, ESI†).

**Table 1** Relative energies of **I**, **II**, **III**, **IV**, and **1a** based on the optimised structures obtained by DFT calculations [(U)B3LYP/6-31G\*]

			T-form (diradical)	
$\Delta E_{\text{rel}}$ [kcal mol <sup>−1</sup> ]	F-form	T-form	Singlet	Triplet
<b>I</b>	0	14.0 <sup>a</sup>	13.6	13.9
<b>II</b>	0	9.40	2.31	2.35
<b>III</b>	0	3.35 <sup>a</sup>	14.3	16.1
<b>IV</b>	0	11.3	12.8	13.4
<b>1a</b>	3.15	0	10.9	13.9

<sup>a</sup> Optimised as a transition-state structure.



## Preparation

The N<sub>4</sub>AQD derivatives **1a–d** were synthesised as follows (Scheme 1). First, the dehydrative condensation of 2,3,5,6-tetraamino-1,4-benzoquinone **2**<sup>34</sup> with benzil to give **3**,<sup>35</sup> and subsequent double dibromo-olefination led to the precursor **4**.<sup>36,37</sup> Next, Suzuki–Miyaura cross-coupling reactions of **4** with the corresponding boronic acid or boroxin afforded the desired N<sub>4</sub>AQD derivatives **1a–d** in 78–99% yields. Reference compounds **I**,<sup>38</sup> **III**,<sup>39</sup> and **IV** were synthesised in a similar manner (Scheme S1, ESI†). These seven AQD derivatives were characterised by <sup>1</sup>H/<sup>13</sup>C NMR and IR spectroscopies, and field desorption (FD)-mass spectrometry (see the ESI†).

## Structure in solution

To gain insight into the solution-phase structure of the N<sub>4</sub>AQD derivatives, the UV/Vis spectrum of the parent phenyl derivative **1a** (R = H) was measured in CH<sub>2</sub>Cl<sub>2</sub>. Compared to the spectra of the reference compounds **I**, **III**, and **IV**, a red shift of the absorption maxima was clearly observed ( $\lambda_{\text{max}}$  = 470 nm for **1a** and  $\lambda_{\text{max}}$   $\approx$  320 nm for **I**, **III**, and **IV**), and the solution of **1a** is reddish-orange whereas those of **I**, **III** and **IV** are colourless to pale yellow (Fig. 2a and Table 2). The HOMO–LUMO gaps of each derivative predicted by DFT calculations were estimated to be 4.11 eV, 3.67 eV, and 3.87 eV for the F-forms of **I**, **III**, and **IV**, respectively, and 2.38 eV for the T-form of **1a**, which is in good agreement with the trend of the absorption end of each derivative. These results suggest that only **1a** can have a contribution from a twisted structure in solution. In addition, the UV/Vis spectra of **1b** (R = CH<sub>3</sub>), **1c** (R = F), and **1d** (R = Cl) in CH<sub>2</sub>Cl<sub>2</sub> were almost the same as that of **1a** (Fig. S15 and Table S7, ESI†), indicating that all N<sub>4</sub>AQD derivatives **1** have contributions from the T-form. According to variable-temperature (VT) <sup>1</sup>H NMR analysis of derivative **1b**, no broadening of signals was observed upon heating from 303 K to 393 K in DMSO-*d*<sub>6</sub> (Fig. S16, ESI†). Thus, there is no contribution of open-shell species. On the other hand, some proton signals shifted upon heating in DMSO-*d*<sub>6</sub> and cooling in CD<sub>2</sub>Cl<sub>2</sub> (Fig. S17, ESI†). To clarify the reason why some proton signals shifted upon elevating the temperature, UV/Vis spectra of **1b** were measured in the range from 293 K to 373 K in DMSO (Fig. 2b). Upon elevating the temperature, a decrease in the absorption band around 500 nm

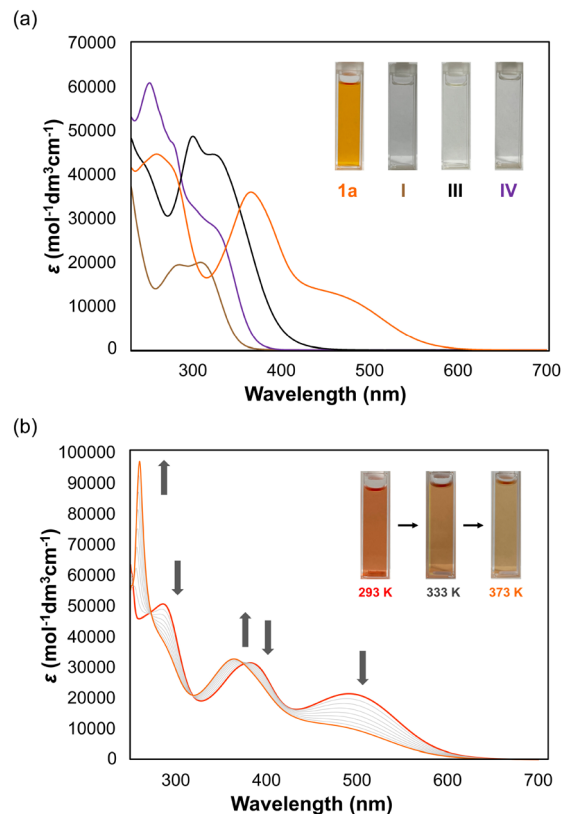
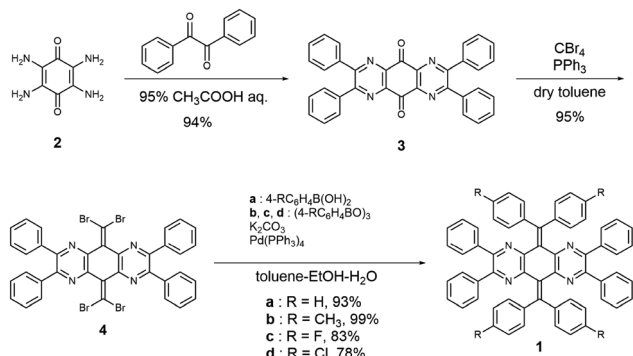


Fig. 2 (a) UV/Vis spectra of **1a**, **I**, **III**, and **IV** in CH<sub>2</sub>Cl<sub>2</sub>. (b) VT UV/Vis spectra of **1b** in DMSO from 293 K to 373 K (every 10 K).

Table 2 Experimental and theoretical absorption properties (in CH<sub>2</sub>Cl<sub>2</sub>, at 293 K) of **1a**, **I**, **III**, and **IV**. Shoulder peaks are marked with "sh".  $\lambda_{\text{end}}$  is the absorption end of each spectrum. The energy gap  $\Delta E_{\text{LUMO-HOMO}}^{\text{DFT}}$  was estimated by DFT calculations (B3LYP/6-31G\*)

	$\lambda_{\text{max}}$ [nm] (log $\epsilon$ )	$\lambda_{\text{end}}$ [nm]	$\Delta E_{\text{LUMO-HOMO}}^{\text{DFT}}$ [eV]/[nm]
<b>1a</b> (Ph <sub>8</sub> N <sub>4</sub> AQD)	sh 470 (4.09) 366 (4.56) 259 (4.65)	657	3.50/354 (F-form) 2.38/520 (T-form)
<b>I</b> (Ph <sub>4</sub> AQD)	308 (4.30) 285 (4.29)	453	4.11/302
<b>III</b> (Ph <sub>4</sub> N <sub>4</sub> AQD)	322 (4.65) 300 (4.69)	498	3.67/338
<b>IV</b> (Ph <sub>8</sub> AQD)	sh 320 (4.46) sh 275 (4.68) 251 (4.78)	409	3.88/319



Scheme 1 Preparation of **1**.

was observed, accompanied by a change in the colour of the solution from reddish-orange to nearly yellow. These observations clearly indicate that the F- and T-forms are in equilibrium in solution. Thus, a thermochromic behaviour was observed for **1b**, since the contribution of T-form is more dominant at lower temperature while that of the metastable F-form increases at higher temperatures according to the Boltzmann distribution. Due to the rapid change in structures, all N<sub>4</sub>AQD derivatives **1** are non-fluorescent in solution. Such fluorescence quenching can often be observed in tetraarylethylenes,<sup>40,41</sup> which are the representative molecules with aggregation-induced emission properties.<sup>42</sup>

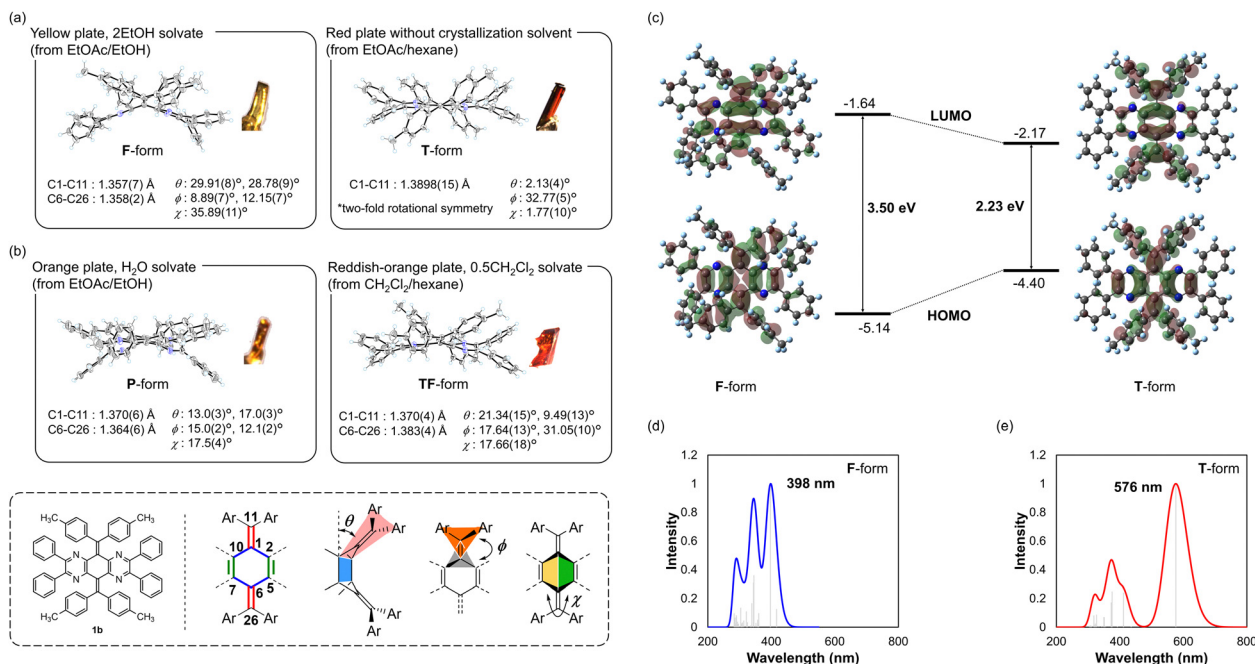


## X-Ray analyses

Since the presence of the T-form was indicated for N<sub>4</sub>AQD derivatives (**1**) in solution, we turned our attention to their structures and photophysical properties in the solid states. X-ray structures were investigated using single crystals of N<sub>4</sub>AQD derivatives **1a–d**, which were obtained by recrystallization from various solvents. For the tolyl derivative **1b**, recrystallization from EtOAc/EtOH produced yellow crystals as EtOH solvate, while that from EtOAc/hexane provided red crystals without containing any solvent. X-Ray analyses of both crystals revealed that **1b** adopts the F- and T-forms in yellow crystals containing EtOH and red crystals, respectively (Fig. 3a). To investigate the origin of their colour, we performed DFT calculations (B3LYP/6-31G\*) using the obtained crystallographic coordinates, and the results indicated that the T-form has a higher HOMO and lower LUMO than the F-form, *i.e.*, the electronic structures of these conformers are significantly different (Fig. 3c and Table S10, ESI<sup>†</sup>). Actually, time-dependent (TD)-DFT calculations (B3LYP/6-31G\*) predicted a *ca.* 180 nm difference in the simulated absorption maxima between these conformers (398 nm for F-form and 576 nm for T-form), which well reproduces the colour of each crystal (Fig. 3d, e and Fig. S21, ESI<sup>†</sup>). On the other hand, the only X-ray structures obtained for the reference compounds **III** and **IV** were folded (Fig. S10, ESI<sup>†</sup>). These results emphasise that both the nitrogen atoms and the *vic*-diphenyl groups introduced into the central AQD skeleton contribute to stabilization of the T-form.

The most important finding is that the obtained crystal structures were not only the F- and T-forms but also several other conformations, which could not be classified as either of them.

For orange-coloured H<sub>2</sub>O-solvated crystal and reddish-orange-coloured CH<sub>2</sub>Cl<sub>2</sub>-solvated crystal of **1b**, the P-form with an almost normal planar geometry of the C=C double bond and the TF-form with an intermediary structure between F- and T-forms were observed, respectively (Fig. 3b and Table S5, ESI<sup>†</sup>). Furthermore, for other derivatives (*e.g.*, phenyl derivative **1a**, fluorophenyl derivative **1c**, and chlorophenyl derivative **1d**), multiple conformations were observed in their crystals, regardless of the difference in the steric and electronic properties of the substituents on the aryl groups (Table 3 and Fig. S6–S9, ESI<sup>†</sup>). Since these intermediate structures are considered to be unstable in solution, their abundance ratio is generally zero. In fact, optimization from the obtained crystallographic coordinates as the initial structure yielded the T-form, which is the most stable structure. Even though there are some intermolecular contacts between molecules in each crystal, no significant interactions were observed to stabilise the specific conformation (Fig. S11–S14, ESI<sup>†</sup>). Based on these results, the intermediate structures observed in the pseudopolymorphs are stabilised by the crystal lattice, which is modified by the solvent molecules filling a void while forming a crystal. Accordingly, we clarified that multiple intermediate structures as well as the F- and T-forms can arise in the same molecule for AQD-type OCEs. Although a few examples have been described previously, where multiple conformations with a slight difference in dihedral angles were observed in (pseudo)polymorphs,<sup>43,44</sup> this study is unique because a drastic change in colour has been achieved for multiple conformers with a large difference in molecular geometry such as the F-, T-, P-, and TF-forms. Thus, specific electronic properties such as HOMO–LUMO gaps based



**Fig. 3** ORTEP drawings of **1b** [(a) F-form in 2EtOH solvate (recrystallised from EtOAc/EtOH) and T-form without any solvent (recrystallised from EtOAc/hexane), and (b) P-form in H<sub>2</sub>O solvate (recrystallised from EtOAc/EtOH) and TF-form in 0.5CH<sub>2</sub>Cl<sub>2</sub> solvate (recrystallised from CH<sub>2</sub>Cl<sub>2</sub>/hexane)]. The solvent molecules are omitted for clarity. (c) HOMO and LUMO levels calculated using the DFT method (B3LYP/6-31G\*). Simulated UV/Vis spectra of (d) the F-form and (e) the T-form obtained by TD-DFT calculations (B3LYP/6-31G\*) based on the crystallographic coordinates.



	Conformation	Colour <sup>b</sup>	Shape	Solvate	Recrystallization solvent
<b>1a</b> (R = H)	<b>F</b> -form	Yellow	Plate	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /hexane
	<b>F</b> -form	Yellow	Plate	CHCl <sub>3</sub>	CHCl <sub>3</sub> /EtOH
	<b>T</b> -form	Red	Block	None	CH <sub>2</sub> Cl <sub>2</sub> /EtOH
<b>1b</b> (R = CH <sub>3</sub> )	<b>F</b> -form	Yellow	Plate	2EtOH	EtOAc/EtOH
	<b>P</b> -form	Orange	Plate	H <sub>2</sub> O	EtOAc/EtOH
	<b>P</b> -form	Orange	Plate	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /hexane
	<b>P</b> -form	Orange	Plate	0.5CHCl <sub>3</sub>	CHCl <sub>3</sub> /EtOH
	<b>TF</b> -form	Reddish-orange	Plate	0.5CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /hexane
	<b>TF</b> -form	Reddish-orange	Block	0.5hexane	EtOAc/hexane
	<b>TF</b> -form	Reddish-orange	Plate	0.5EtOAc	EtOAc/EtOH
	<b>TF</b> -form	Reddish-orange	Plate	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /EtOH
	<b>T</b> -form	Red	Plate	None	EtOAc/hexane
	<b>F</b> -form	Yellow	Block	EtOAc	EtOAc
<b>1c</b> (R = F)	<b>F</b> -form	Yellow	Plate	0.25CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /hexane
	<b>F</b> -form	Yellow	Plate	1.25CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /EtOH
	<b>F</b> -form	Yellow	Plate	2CHCl <sub>3</sub>	CHCl <sub>3</sub> /EtOH
	<b>F</b> -form + semi- <b>P</b> -form <sup>a</sup>	Orange	Plate	H <sub>2</sub> O	CHCl <sub>3</sub> /hexane
	<b>TP</b> -form	Red	Needle	CHCl <sub>3</sub>	CHCl <sub>3</sub> /hexane
	<b>TP</b> -form <sup>a</sup>	Red	Needle	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /hexane
	<b>TP</b> -form <sup>a</sup>	Red	Needle	EtOH	EtOAc/EtOH
	<b>TF</b> -form	Reddish-orange	Plate	CHCl <sub>3</sub>	CHCl <sub>3</sub> /hexane
	<b>P</b> -form <sup>a</sup>	Orange	Plate	CHCl <sub>3</sub>	CHCl <sub>3</sub> /EtOH
	<b>P</b> -form	Orange	Plate	CH <sub>2</sub> Cl <sub>2</sub>	CH <sub>2</sub> Cl <sub>2</sub> /hexane
<b>1d</b> (R = Cl)	<b>TF</b> -form	Reddish-orange	Plate	CHCl <sub>3</sub>	CHCl <sub>3</sub> /hexane
	<b>T</b> -form	Red	Needle	None	CHCl <sub>3</sub> /hexane

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**Fig. 4** (a) Emission spectra of each conformation of **1c** in the solid state. (b) ORTEP drawings of **1c** [F-form in EtOAc solvate (recrystallised from EtOAc), TP-form in CH<sub>2</sub>Cl<sub>2</sub> solvate (recrystallised from CH<sub>2</sub>Cl<sub>2</sub>/hexane), TF-form in CHCl<sub>3</sub> solvate (recrystallised from CHCl<sub>3</sub>/hexane), and those in which the F- and semi-P-forms exist as two crystallographically independent molecules in H<sub>2</sub>O solvate (recrystallised from CHCl<sub>3</sub>/hexane)]. Each of them was classified by the difference in dihedral angle  $\theta$ ,  $\phi$ , and  $\chi$  for which emission spectra were measured. The solvent molecules are omitted for clarity. (c) Luminescence properties and dihedral angle  $\phi$  of each conformation of **1c** were recorded in the solid state. All crystals were excited at 380 nm. Emission lifetimes were detected at the longer emission maximum.  $\lambda_{ex}$  is the value of the peak in the excitation spectra of **1c**.

powder with EtOAc followed by drying, which indicates that the HOMO–LUMO gap can be modulated based on the conformations by being exposed to external stimuli. To investigate the origin of the change in the photophysical properties upon mechanical stimulation, powder X-ray diffraction (PXRD) analyses were conducted for each solid-state sample of **1c**. As shown in Fig. 5c, since the PXRD pattern of the ground sample (orange line) has no sharp peaks, a phase transition from crystal to amorphous occurred upon grinding the pristine crystal. As a result, the metastable F-form could no longer be maintained in



**Fig. 5** (a) Photographs of **1c** under room light (upper) and UV light (lower) in various states. (b) Changes in the emission spectrum for **1c**. (c) Changes in the PXRD pattern of **1c**. The black line is the simulation pattern from single-crystal X-ray diffraction (SCXRD).

an amorphous state, and thus the more stable twisted conformation appeared. Next, the sharp diffraction pattern was regenerated by treating the sample with EtOAc followed by drying (Fig. 5c, from the orange line to the grey line), which demonstrates that **1c** exhibits mechanochromism in the ground and excited states based on the crystal-to-amorphous phase transition. This phenomenon was realised due to the “flexibility” granted to the AQD-type OCE, *i.e.*, not only the reduction of steric hindrance by the introduction of nitrogen atoms into the central framework but also the introduction of *vic*-diphenyl groups on the fused pyrazine rings are important for creating multiple conformations with tunable photophysical properties.

## Conclusions

In summary, we have designed and synthesised four kinds of tetraaza AQD derivative **1** with moderately reduced steric hindrance in the overcrowded fjord region with the introduction of *vic*-diphenyl groups. We succeeded in observing not only the F- and T-forms but also many other conformers, *e.g.*, the P-form, TF-form, and TP-form, in pseudopolymorphs of **1**, and this is the first demonstration of such multiple conformations for AQD-type OCEs. In addition, we demonstrated that the HOMO/LUMO levels can be modulated by the conformations in crystals. Thus, the changes in colour tone and emission of the crystals were accounted for by the difference in the electronic configurations of the conformers, and these quinodimethanes should be candidates for the development of advanced materials exhibiting conformation-dependent multi-chromism. It is particularly significant that the formation of twisted conformations with long-wavelength absorption/NIR-emission was achieved without the involvement of charge-separation states. This study has realised the control of absorption and luminescence properties derived from conformational changes, which



is a very rare phenomenon in organic molecules, and revealed the potential value of novel sensor materials composed of simple but flexible organic molecules exhibiting an unprecedented gradual change in colour.

## Conflicts of interest

There are no conflicts to declare.

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