

PAPER

[View Article Online](#)
[View Journal](#) | [View Issue](#)Cite this: *Dalton Trans.*, 2021, **50**,
13264Highly conductive and stable Co₉S₈ thin films by
atomic layer deposition: from process develop-
ment and film characterization to selective and
epitaxial growth†Miika Mattinen,^{‡a} Timo Hatanpää,^{‡a} Kenichiro Mizohata,^{‡b} Jyrki Räisänen,^b
Markku Leskelä^{‡a} and Mikko Ritala^{‡a*}

Co₉S₈ is an interesting sulfide material with metallic conductivity that has shown promise for various energy applications. Herein, we report a new atomic layer deposition process producing crystalline, pure, and highly conductive Co₉S₈ thin films using CoCl₂(TMEDA) (TMEDA = *N,N,N',N'*-tetramethylethylenediamine) and H₂S as precursors at 180–300 °C. The lowest resistivity of 80 μΩ cm, best uniformity, and highest growth rate are achieved at 275 °C. Area-selective deposition is enabled by inherent substrate-dependency of film nucleation. We show that a continuous and conductive Co₉S₈ film can be prepared on oxide-covered silicon without any growth on Si–H. Besides silicon, Co₉S₈ films can be grown on a variety of substrates. The first example of an epitaxial Co₉S₈ film is shown using a GaN substrate. The Co₉S₈ films are stable up to 750 °C in N₂, 400 °C in forming gas, and 225 °C in O₂ atmosphere. The reported ALD process offers a scalable and cost-effective route to high-quality Co₉S₈ films, which are of interest for applications ranging from electrocatalysis and rechargeable batteries to metal barrier and liner layers in microelectronics and beyond.

Received 12th July 2021,
Accepted 20th August 2021

DOI: 10.1039/d1dt02315b

rsc.li/dalton

Introduction

Cobalt forms a range of crystalline sulfides, including cubic Co₄S₃, cubic Co₉S₈ (mineral name cobalt pentlandite), hexagonal CoS (jaipurite), cubic Co₃S₄ (linnaeite), and cubic CoS₂ (cattierite).^{1–3} Co₉S₈ displays metallic conductivity and it is one of the most studied and stable cobalt sulfide phases. Co₉S₈ has been actively explored as an electrode for electrocatalytic water splitting including both oxygen evolution reaction (OER)^{4–6} and hydrogen evolution reaction (HER),^{5,7,8} as well as for lithium ion batteries,⁹ zinc-air batteries,⁶ dye-sensitized solar cells,^{10–12} and supercapacitors.^{13,14} Due to its high conductivity and stability, Co₉S₈ may also be of interest as a barrier and liner layer for interconnects in microelectronics, provided that ultrathin continuous films can be prepared. This

possibility seems particularly interesting in light of Intel's introduction of cobalt interconnects at the 10 nm node.¹⁵

To date, several methods have been used to deposit Co₉S₈ nanoparticles, including colloidal synthesis,^{5,11} solvothermal¹⁶ and hydrothermal synthesis,⁴ thermal decomposition,¹⁷ and spray pyrolysis.¹⁸ Rough thin films from hundreds of nanometers to micrometers in thickness have been prepared by chemical vapor deposition (CVD),^{19–21} electrodeposition,^{9,12} and solvothermal synthesis.^{7,10}

The aforementioned methods have, however, struggled in producing thin, continuous Co₉S₈ thin films with accurately controlled, uniform thickness and high material quality. To overcome these limitations, we use atomic layer deposition (ALD), which is an advanced modification of CVD. In ALD, alternately pulsed precursors react with the substrate one-by-one in self-limiting manner, while the reactor is purged between the precursor pulses to ensure that the precursors do not react with each other in the gas phase.^{22–24} Each ALD cycle deposits a (sub)monolayer of material, which gives ALD unmatched control over film thickness, uniformity, and conformality on both large and three-dimensional substrates. Excellent scalability, reproducibility, and relatively low deposition temperatures from room temperature to about 500 °C are also important characteristics of ALD. The unique advantages of ALD have led to its widespread industrial use.^{25,26} ALD

^aDepartment of Chemistry, University of Helsinki, P.O. Box 55, FI-00014, Finland.
E-mail: mikko.ritala@helsinki.fi

^bDepartment of Physics, University of Helsinki, P.O. Box 43, FI-00014, Finland

†Electronic supplementary information (ESI) available: Results of additional deposition experiments and film characterization, further details and discussion of epitaxial growth, XRD measurements on film stability in different atmospheres. See DOI: 10.1039/d1dt02315b

‡Present address: Department of Applied Physics, Eindhoven University of Technology, P.O. Box 513, 5600 MB Eindhoven, The Netherlands.



can be used to deposit numerous materials,²⁷ including elements, oxides, and nitrides, as well as many sulfides.²⁸ However, when developing an ALD process for a new material, it is often a challenge to identify molecules that fulfill the demanding requirements set for an ALD precursor, including volatility, high reactivity, and lack of self-decomposition pathways at the deposition temperature.

So far, only one ALD process exists for Co₉S₈, which uses Co(ⁱPrAMD)₂ and H₂S at 100–225 °C.^{6,8,13,14,29} In addition, the same Co(ⁱPrAMD)₂ precursor has been used with H₂S plasma to deposit CoS₂ thin films at 80–225 °C.³⁰ However, Co(ⁱPrAMD)₂ is an expensive precursor (~200 eur g⁻¹),³¹ which limits its use for large-scale applications. Furthermore, the resistivity of the Co₉S₈ films deposited by this process is relatively high, 3–5 mΩ cm.¹³

CoCl₂(TMEDA) (TMEDA = *N,N,N',N'*-tetramethylethylenediamine) was recently introduced as an affordable alternative to existing cobalt ALD precursors. Its favorable ALD precursor characteristics have been shown for deposition of high-quality CoO (ref. 32) and Co₃Sn₂ (ref. 33 and 34) thin films. In this article, CoCl₂(TMEDA) is used with H₂S to deposit crystalline Co₉S₈ films with a low resistivity on the order of 100 μΩ cm. Detailed investigation of the ALD process at temperatures ranging from 180 to 300 °C is followed by characterization of morphology, crystallinity, resistivity, and composition of the deposited films. Film nucleation and growth on different substrates is also studied in detail to assess the potential of the process for area-selective and epitaxial growth as well as compatibility with different substrates. Finally, high-temperature XRD measurements are used to evaluate the stability and applicability of the films in harsh conditions.

Experimental

Film deposition

Cobalt sulfide thin films were deposited using a commercial, hot-wall, cross-flow ALD reactor (F120, ASM Microchemistry) operated under approximately 5 mbar pressure.³⁵ Nitrogen (N₂, 99.999%, AGA) was further purified using a SAES Microtorr MC1-902F purifier and was used as the carrier and purge gas at a flow rate of 400 sccm. Dichloro(*N,N,N',N'*-tetramethylethylenediamine)cobalt(II) [CoCl₂(TMEDA)] precursor was synthesized in house using a simple one-step room temperature route giving 99% yield.³² The precursor was heated to 170 °C in an open boat held inside the ALD reactor and pulsed by inert gas valving. The precursor is a solid at room temperature, but it melts under the source conditions used in this work. The mildly air sensitive precursor was stored and handled in a N₂ glove box before transferring to the ALD reactor.

The sulfur precursor hydrogen sulfide (H₂S, 99.5%, Linde) was further purified using a SAES Microtorr MC1-302F purifier. The flow rate of H₂S was set to 14 sccm under a continuous flow using a mass flow meter and a needle valve and it was pulsed into the reactor using an external solenoid valve.

Caution! Safe use of highly toxic and flammable H₂S gas requires a properly designed ALD reactor and laboratory space.

The H₂S bottle was stored in a ventilated gas cabinet and the H₂S lines were designed to be compatible with H₂S, using VCR metal and EPDM polymer seals. The reactor exhaust was bubbled through aqueous Cu(NO₃)₂ solution to remove H₂S downstream of the vacuum pump. An example of ALD reactor modifications for H₂S has been published by Dasgupta *et al.*³⁶

Cobalt sulfide films were mostly deposited on 5 × 5 cm² Si (100) and soda lime glass (SLG) substrates. The Si substrates were used as supplied with a native oxide layer. The SLG substrates were cleaned using successive ultrasonic baths of an alkaline detergent (Industrial Strength Cleaner, Branson), tap water, ethanol, and deionized water (10 minutes each, room temperature) followed by careful rinsing using deionized water and 50 : 50 (v/v%) deionized water/ethanol solution and blown dry using pressurized air. Smaller pieces of other substrates were used for selected depositions, including thermal SiO₂ (300 nm)/Si, Si-H(100), sapphire (2" wafer, C-plane off to M-plane by 0.2 ± 0.1°, double-side polished, University Wafer), muscovite mica (15 mm disks, V1 grade, Nano-Tec), GaN (epitaxially grown on Si(111) by MOCVD with an AlN nucleation layer and AlGaN buffer layers at Imec, see Fig. S21 in the ESI†), and different ALD films (Al₂O₃, Ir, and TiO₂) prepared in-house on Si. Muscovite mica was cleaved using a double-edged razor blade immediately before deposition to reveal a fresh (001) surface.³⁷ Sapphire wafers were rinsed with ethanol, acetone, and isopropanol followed by annealing in air at 1000 °C for 2 h to create a surface consisting of flat (0001) terraces.³⁸ Si-H was prepared by dipping a silicon(100) substrate with a native oxide into a 1 v% aqueous HF solution for 2 min followed by rinsing with deionized H₂O, blowing dry with pressurized N₂, and loading into the reactor. The other substrates were stored in air and blown clean of particles using pressurized N₂.

Film characterization

Film morphology was examined by scanning electron microscopy (SEM, Hitachi S-4800) and atomic force microscopy (AFM, Veeco Multimode V). Tapping mode AFM imaging was performed in air using silicon probes with a nominal radius of less than 10 nm (Bruker). The images were flattened with no other image processing performed. Roughness was calculated as a root-mean-square (*R_q*) value.

Film thicknesses were measured by energy-dispersive X-ray spectrometry (EDS, Oxford INCA 350 connected to the Hitachi S-4800 SEM). GMRFilm software³⁹ was used to convert the measured Co and S Kα *k*-ratios to film thicknesses assuming bulk density of Co₉S₈ (5.34 g cm⁻³).⁴⁰ X-ray reflectivity (XRR, Rigaku SmartLab) was also used for thickness measurements. XRR was mostly used for films thinner than approximately 15 nm and EDS for films thicker than 15 nm. Reliable use of EDS for thinner films was prevented by low signal intensity, whereas the high roughness of thicker films prevented accurate XRR measurements of them.

EDS was also used to measure the S/Co ratio of the films, while complete elemental composition was analyzed by time-of-flight elastic recoil detection analysis (ToF-ERDA) using 40



MeV $^{127}\text{I}^{7+}$ ion beam. The angle between the incident beam and sample surface was 16° and that between the recoiled beam and sample surface was 24° , resulting in 40° incident beam to detector angle. The film surface and film/substrate interfaces were excluded when evaluating the compositions from the ToF-ERDA results.

Sheet resistance was measured by a four-point-probe (CPS Probe station connected to a Keithley 2400 SourceMeter). The sheet resistance was converted to resistivity using the thicknesses measured by EDS or XRR.

Crystallinity was studied by X-ray diffraction (XRD, Rigaku SmartLab) in either grazing incidence (GIXRD, $\omega = 1^\circ$) or $\theta - 2\theta$ geometry. In addition, rocking curve (ω scan) measurements were used to quantify out-of-plane texture. Epitaxial in-plane registration of the films grown on mica and GaN was studied using in-plane pole figure and in-plane ϕ scan (incident angle 0.3°) measurements with the diffraction angle (denoted $2\theta_B$ and $2\theta_\chi$ in these measurements, respectively) fixed to known substrate and film reflections.^{41,42} The diffractometer was equipped with an in-plane arm and a non-monochromatized copper X-ray tube mainly emitting Cu K α radiation ($\lambda \approx 1.54$ Å, consisting of Cu K α_1 and K α_2 at 1.5406 and 1.5443 Å). Minor contributions from Cu K β ($\lambda \approx 1.39$ Å) and W L α ($\lambda \approx 1.48$ Å) lines were also present, which is apparent especially in the strongest substrate reflections. A Rigaku D/Tex Ultra 250 detector was used to suppress the high fluorescent background originating from cobalt in the films.

Thermodynamic calculations and visualization of epitaxy

Thermodynamic calculations of deposition chemistry were done using Outokumpu HSC Chemistry for Windows software (version 5.11).

VESTA software⁴³ was used to visualize the orientation of Co_9S_8 on GaN using data from the Inorganic Crystal Structure Database (ICSD; Co_9S_8 no. 23929, GaN no. 54698).

High-temperature X-ray diffraction

In situ high-temperature XRD (HTXRD) measurements were performed using an Anton-Paar HTK1200N oven connected to a PANalytical X'Pert Pro MPD diffractometer employing a non-monochromatized copper X-ray tube in grazing incidence ($\omega = 1^\circ$) geometry. For each measurement, either an laboratory air (1 atm), O_2 (1 atm, AGA, 99.999%), N_2 (30 mbar achieved by a diaphragm pump, AGA, 99.999%, further purified using an Entegris Gatekeeper purifier), or forming gas (1 atm, 10% H_2 in N_2 , AGA) or dynamic vacuum (approximately 10^{-5} mbar achieved by a turbomolecular pump backed by a diaphragm pump) was used. The studied Co_9S_8 films (deposited on Si (100) at 275°C using 1000 cycles) were first heated from room temperature up to the starting temperature of 100 or 125°C at a heating rate of $10^\circ\text{C min}^{-1}$. Thereafter, a diffractogram was recorded every 25°C , which took 27 min at each temperature, before heating to the next temperature at a heating rate of $10^\circ\text{C min}^{-1}$. The measurements were continued up to 725°C or 750°C , resulting in a total measurement time of approximately 15 h per experiment.

Results and discussion

Film deposition

Cobalt sulfide thin films were deposited by ALD using alternating pulses of $\text{CoCl}_2(\text{TMEDA})$ and H_2S separated by nitrogen purge steps. Initial deposition studies showed that Co_9S_8 films could be deposited in the whole available temperature range of 180 to 300°C . The lower temperature limit was dictated by the $\text{CoCl}_2(\text{TMEDA})$ source temperature of 170°C and the higher limit by the onset of $\text{CoCl}_2(\text{TMEDA})$ decomposition.³² The neutral TMEDA adduct ligand makes the non-volatile cobalt(II) chloride volatile, while it has little effect on the deposition chemistry.^{32,34} Thermodynamically, the reactions between CoCl_2 and H_2S producing cobalt sulfide in different possible stoichiometries (Co_9S_8 , CoS , Co_3S_4 , and CoS_2) are favorable with similar ΔG values of approximately -125 to -145 kJ mol^{-1} of CoCl_2 at 180 – 300°C (section S1 in the ESI†). Thus, the corresponding ALD surface reactions of our process are also likely to be very favorable.

Next, the self-limiting nature of the surface reactions, a unique characteristic of ALD, was assessed. With increasing $\text{CoCl}_2(\text{TMEDA})$ pulse time, the growth rate appeared to saturate at 180°C to approximately 0.2 Å per cycle when a pulse length of at least 1.0 s was used (Fig. 1a). At 250°C , the growth rate was doubled to approximately 0.4 Å per cycle using pulse lengths of 0.5 to 2.0 s. At 275°C , the growth rate was clearly higher, for example 0.7 Å per cycle using 1.0 s $\text{CoCl}_2(\text{TMEDA})$ pulses. Increasing the $\text{CoCl}_2(\text{TMEDA})$ pulse length at 275°C led to an increase of the growth rate, reaching 1.0 Å per cycle using 4.0 s pulses. No complete saturation was observed even with 8.0 s pulses, although the increase in the growth rate slowed down with the longest pulses tested (ESI, Fig. S2†). However, no film growth was observed when only $\text{CoCl}_2(\text{TMEDA})$ was pulsed repeatedly at 275°C on silicon or on an ALD-grown cobalt sulfide film (ESI, Fig. S3†), ruling out decomposition of $\text{CoCl}_2(\text{TMEDA})$.

Supported by observed changes in crystalline texture as a function of $\text{CoCl}_2(\text{TMEDA})$ pulse length, we propose that the slow saturation at 275°C may be due to different crystalline facets having different growth rates as well as kinetics (see section S3 in the ESI† for the full explanation). Despite the atypical saturation, conformal films were deposited on trench structures at 275°C (ESI, Fig. S4†).

The film uniformity improved at 275°C when long $\text{CoCl}_2(\text{TMEDA})$ pulse lengths were used. The non-uniformity of sheet resistance over the 5×5 cm^2 silicon substrate decreased from approximately 20% using ≤ 1 s pulses to less than 5% using 4 s $\text{CoCl}_2(\text{TMEDA})$ pulses. At 180°C , a similar trend with slightly higher non-uniformities was observed, whereas at 250°C non-uniformity was high, 30–50%, regardless of the pulse lengths applied (Fig. 1b). In addition, the film resistivity was the lowest at 275°C , namely 80 – 120 $\mu\Omega$ cm. The S/Co atomic ratio was measured to be 0.9 – 1.0 for the majority of the films (Fig. 1c) in agreement with the Co_9S_8 phase detected by XRD, which has a theoretical S/Co ratio of 0.89 . The same Co_9S_8 phase has also been reported for the $\text{Co}(\text{PrAMD})_2 + \text{H}_2\text{S}$



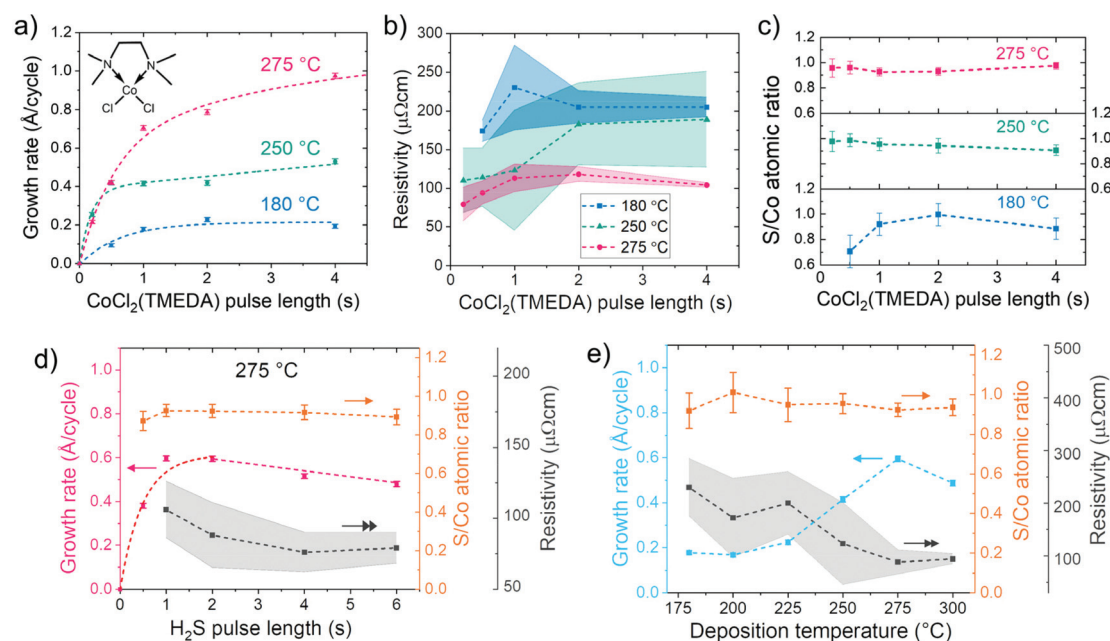


Fig. 1 (a) Growth rate, (b) resistivity, and (c) S/Co atomic ratio versus CoCl₂(TMEDA) pulse length at different deposition temperatures. Schematic structure of the CoCl₂(TMEDA) precursor is shown as an inset in (a). Growth rate, resistivity and S/Co atomic ratio versus (d) H₂S pulse length (at 275 °C) and (e) deposition temperature. Thicknesses and S/Co ratios were determined by EDS. The error bars in growth rates and S/Co ratios represent statistical measurement uncertainty (one standard deviation). The shaded areas around resistivity values represent film uniformity, *i.e.* variation (one standard deviation) in sheet resistance over the 5 × 5 cm² silicon substrate (resistivities calculated using a single thickness value per sample). The dashed lines are meant to guide the eye. The films were deposited on silicon (native oxide) using, unless otherwise noted, 1000 cycles with 1.0 s CoCl₂(TMEDA) and 2.0 s H₂S pulses separated by 1.0 s N₂ purges.

process.^{6,8,13,14,29} Deposition temperature of 275 °C was chosen for further experiments, unless otherwise noted, as it combined the highest growth rate with the best uniformity and lowest resistivity.

Varying the H₂S pulse length at 275 °C showed that a growth rate of 0.6 Å per cycle was reached using 1.0 s H₂S pulses, when the CoCl₂(TMEDA) pulse was fixed to 1.0 s. The growth rate slightly decreased from 0.6 to 0.5 Å per cycle with an increase of the H₂S pulse length from 1.0 to 6.0 s (Fig. 1d). The decrease in growth rate is postulated to be linked to changes in crystalline texture in the opposite direction compared to the changes observed as a function of CoCl₂(TMEDA) pulse (section S3 in the ESI†). With increasing H₂S pulse length, film resistivity decreased from approximately 100 to 80 μΩ cm. The S/Co ratio of the films remained unchanged within the measurement accuracy at approximately 0.9 when the H₂S pulse length was varied.

One second N₂ purges between the precursor pulses were found to be sufficient for removing excess precursor and reaction byproducts, thus preventing any gas phase reactions (ESI, Fig. S5†). Most of the films discussed later in this article were deposited using a pulse sequence optimized for throughput and film uniformity, namely 1.0 s CoCl₂(TMEDA) and 2.0 s H₂S pulses separated by 1.0 s N₂ purges.

A more systematic look at the effect of deposition temperature confirmed the trends discussed above, *i.e.* an increase in growth rate, a decrease in resistivity, and an improvement in uniformity with an increase of temperature from 180 to 275 °C

(Fig. 1e). At 300 °C, the growth rate started to decrease and decomposition of CoCl₂(TMEDA) was visually observed as a dark film deposited in the quartz parts of the reactor in line with previous studies.³² However, good uniformity and low resistivity were still achieved at 300 °C. The growth rate of our process at 275 °C is higher compared to the previous Co (iPrAMD)₂ + H₂S process with reported growth rates of 0.27–0.4 Å per cycle at 80–165 °C depending on the reactor used.^{6,13,14,29}

The resistivity of the films on silicon and soda lime glass substrates (80–220 μΩ cm depending on deposition conditions) was more than an order of magnitude lower compared to the ALD Co₉S₈ films grown using Co(iPrAMD)₂ and H₂S (3–5 mΩ cm).¹³ The bulk resistivity of metallic Co is 5.6 μΩ cm,⁴⁴ *i.e.* about an order of magnitude lower compared to our Co₉S₈ films. However, the resistivities of ALD Co films reported in the literature ranges from 10 to 200 μΩ cm.⁴⁵ Thus, the low resistivity of our Co₉S₈ films is comparable to even some metallic cobalt ALD films.

Characterization of film crystallinity, morphology, and composition

The films deposited at all of the examined temperatures from 180 to 300 °C consisted of purely cubic Co₉S₈ (cobalt pentlandite) according to XRD (Fig. 2a; International Centre for Diffraction Data Powder Diffraction File (PDF) 86-2273). $\theta - 2\theta$ XRD measurements revealed the films to exhibit preferred (111) orientation or texture (Fig. S10 in the ESI†).



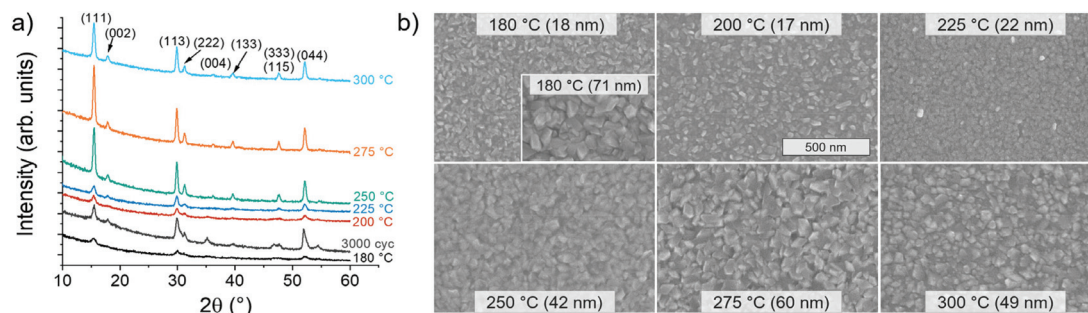


Fig. 2 (a) Grazing incidence X-ray diffractograms (all Miller indices refer to Co_9S_8) and (b) SEM images of films deposited at different temperatures. The films were deposited on silicon (native oxide) using 1000 cycles except for the thicker film deposited at 180 °C using 3000 cycles that is shown in the inset. Film thicknesses determined by EDS are indicated in (b).

Effect of deposition temperature on film crystallinity and texture was evaluated with the aid of a thicker film deposited at 180 °C using 3000 cycles. The broader and lower peaks observed in this film compared to the films of lower thickness deposited at 250 °C and above confirm that the film crystallinity improved and (111) texture strengthened with increasing deposition temperature.

All of the films appeared relatively rough in SEM (Fig. 2b). No major changes in the shape and size of the surface features as a function of deposition temperature were observed for films of similar thickness. The width of the grains ranged from about 50 to 100 nm for 50–70 nm thick films, while in thinner films the grains were smaller as expected.

Film composition was studied by time-of-flight elastic recoil detection analysis (ToF-ERDA), which confirmed the S/Co ratio of 0.9 for the films deposited at 250–300 °C, whereas it was 1.0 for the film deposited at 180 °C (Table 1). Detection of chlorine impurities in the films by ToF-ERDA was complicated by the similar atomic masses of sulfur and chlorine and the limited mass resolution of the setup. Thus, only upper limits of chlorine contents are reported. They suggest a higher chlorine content in the film deposited at 180 °C compared to the films grown at higher temperatures. Low oxygen, carbon, nitrogen, and hydrogen contents of less than 3 at% in total (~4% at 300 °C) were observed in the bulk of the films. Oxygen likely originates from post-deposition oxidation, as supported by an increased concentration of oxygen observed on the film surface (ESI, Fig. S11†), whereas the other impurity elements likely come from the precursors used.

Film nucleation and implications for area-selective deposition

In order to obtain more information on film nucleation, formation of continuous films, and possibilities for area-selective ALD, a series of thinner Co_9S_8 films was prepared on four sub-

strates that were selected based on preliminary experiments. In terms of film thickness, the growth proceeded in a rather similar manner on silicon with either a native oxide (denoted native SiO_2) or a thermal SiO_2 layer (denoted thermal SiO_2) as well as soda lime glass (denoted SLG) as shown in Fig. 3a. During the first 25 cycles, substrate-enhanced growth, in other words a higher growth rate of up to 3 Å per cycle was observed on these substrates, before the growth rate slowed down to the steady-state value of approximately 0.6 Å per cycle at 275 °C (ESI, Fig. S12†). In stark contrast, a clear nucleation delay was observed on HF-dipped hydrogen-terminated silicon (denoted Si-H), on which no growth was seen when 50 cycles or less were applied.

Despite the similar growth rate evolution on native SiO_2 , thermal SiO_2 , and SLG, differences in morphology, continuity, and resistivity were observed. The films deposited on native SiO_2 were electrically conductive already after 10 cycles, at a thickness of approximately 3.3 nm, although there may still be some holes in the film (Fig. 3b; information on additional samples in Fig. S13 and S14 in the ESI†). In contrast, the thinnest conductive film on SLG was 5.2 nm thick (25 cycles) and that on thermal SiO_2 was 7.0 nm thick (50 cycles). The differences in the formation of a continuous film suggest more rapid nucleation and/or lower degree of agglomeration of the initially formed nuclei on native SiO_2 compared to thermal SiO_2 and SLG. Faster nucleation on native SiO_2 would imply a higher number of reactive sites, which is plausible considering that the high temperatures used to prepare thermal SiO_2 are known to decrease the density of reactive –OH groups.⁴⁶ Differences in agglomeration, on the other hand, may be caused by a multitude of factors, both kinetic (e.g. diffusion rate) and thermodynamic (e.g. surface energy) in nature.

Table 1 Elemental composition of Co_9S_8 films deposited on silicon (native oxide) at different temperatures as analyzed by ToF-ERDA. Interfaces with the substrate and air were excluded from the analysis. 1000 cycles were applied except for the 180 °C sample grown using 3000 cycles

Deposition temperature (°C)	Co (at%)	S (at%)	S/Co ratio	Cl (at%)	O (at%)	C (at%)	N (at%)	H (at%)
180	47.0 ± 0.5	48.0 ± 0.5	1.02 ± 0.01	<2.7	0.17 ± 0.01	0.80 ± 0.07	0.20 ± 0.04	1.24 ± 0.09
250	50.1 ± 0.9	45.6 ± 0.9	0.91 ± 0.01	<1.2	0.68 ± 0.06	0.63 ± 0.09	0.19 ± 0.06	1.3 ± 0.2
275	49.4 ± 0.4	45.2 ± 0.4	0.92 ± 0.01	<0.8	0.45 ± 0.04	0.63 ± 0.09	0.12 ± 0.04	1.2 ± 0.2
300	49.9 ± 0.9	45.4 ± 0.9	0.91 ± 0.01	<0.8	0.66 ± 0.06	1.6 ± 0.2	0.29 ± 0.08	1.3 ± 0.2



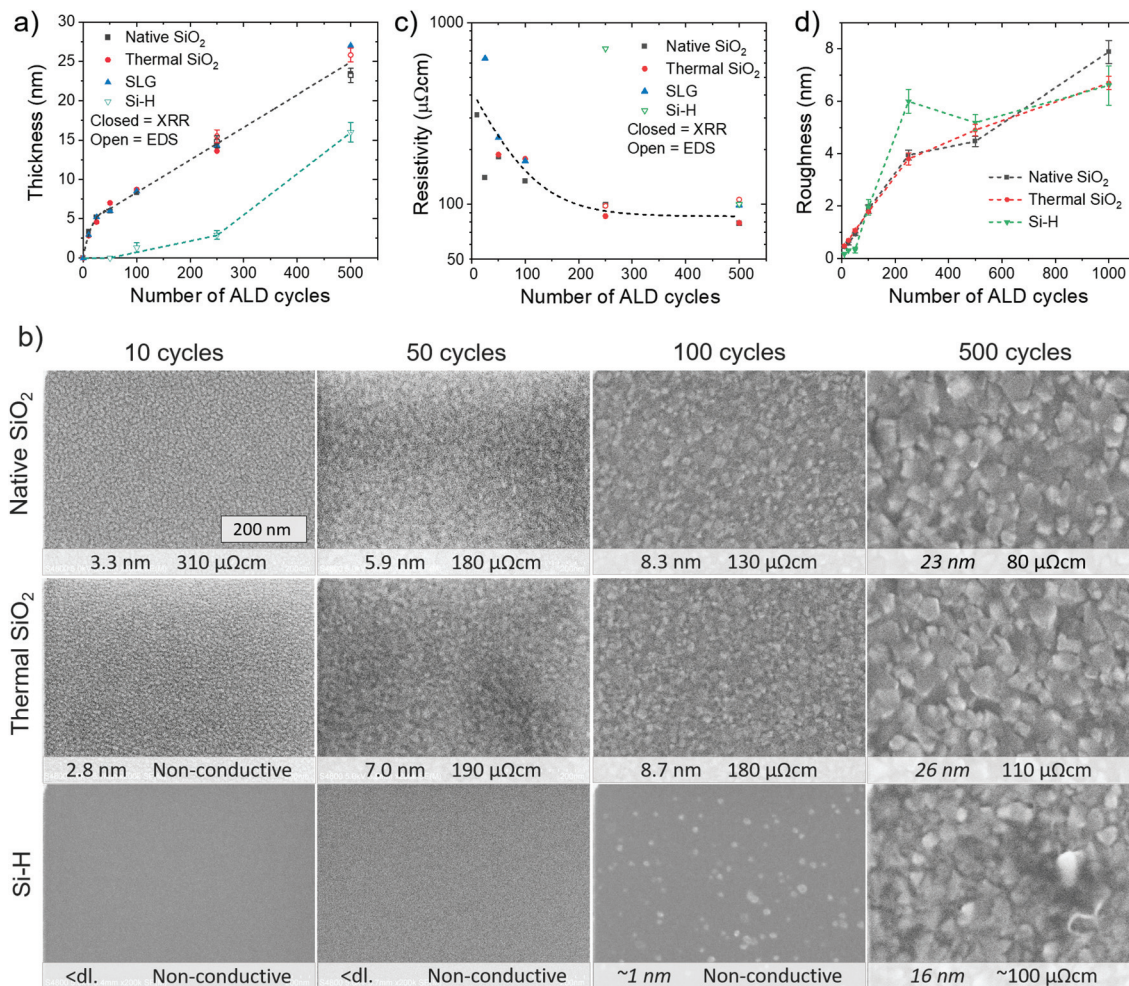


Fig. 3 (a) Thickness, (b) SEM images (thickness and resistivity values also indicated; thicknesses in *italic* measured by EDS, others by XRR; dl. = detection limit), (c) resistivity, and (d) root-mean-square roughness (R_q) of films deposited at 275 °C on different substrates *versus* number of ALD cycles. The dashed lines are meant to guide the eye. The error bars of EDS thicknesses represent statistical measurement uncertainty (one standard deviation). For XRR, the uncertainty was estimated to be smaller than the symbol size. The error bars in roughness represent variation within 3–5 AFM images recorded per sample (one standard deviation).

The resistivities of the ultrathin Co₉S₈ films were expectedly higher compared to the thicker films, up to a few hundred μΩ cm (Fig. 3c). The increased resistivity of the thinner films is likely due to increased electron scattering at the interfaces, as well as at grain boundaries.⁴⁷ The resistivities of films of a given thickness on native SiO₂ were somewhat lower compared to the other substrates (ESI, Fig. S15†). The native SiO₂ substrates had sheet resistances of at least ten times higher than the thinnest Co₉S₈ films, so the substrate conductivity should not have a major effect on the measured values. The lower film resistivity on native SiO₂ is thus likely linked to the faster formation of continuous films as discussed above. For the films deposited using at least 250 cycles (approximately 15 nm thick), the resistivities on the different substrates approached a stable value of less than 100 μΩ cm.

XRD peaks indicating formation of crystalline Co₉S₈ were observed after only 25 cycles on native SiO₂ (ESI, Fig. S16a†). Thus, it seems likely that Co₉S₈ grows in crystalline form from

the first cycles. No major changes in the preferred orientation with increasing film thickness were observed (ESI, Fig. S16b†). Qualitatively, the Co₉S₈ out-of-plane crystallite size on native SiO₂ increased with increasing thickness as shown by decreasing XRD peak widths. At the same time, also the in-plane size of the surface features (grains) increased as observed by SEM (Fig. 3b).

On Si-H, a continuous film with a high resistivity of more than 600 μΩ cm was obtained only after 250 ALD cycles. The nucleation delay on Si-H was confirmed by AFM, which showed practically no change in roughness up to 50 cycles, in contrast to a monotonous increase of film roughness on native and thermal SiO₂ (Fig. 3d, images in Fig. S17 in the ESI†). For thicker films, the roughnesses on native SiO₂, thermal SiO₂, and Si-H were similar and relatively large (4–5 and 6–8 nm using 500 and 1000 cycles). The resistivity and non-uniformity of the films on Si-H remained higher compared to the SiO₂ substrates.



These results indicate that the present ALD process enables inherent area-selective deposition of Co_9S_8 on oxide-covered silicon with readily prepared Si-H as a non-growth surface. The 50 cycle nucleation delay on Si-H means that conductive and continuous, 6–7 nm thick films can be deposited on SiO_2 before any growth is detected on Si-H. Because of the high conductivity of Co_9S_8 , the system may be described as selective growth of a metal (Co_9S_8) on a dielectric (SiO_2) with a semiconductor (Si-H) as the non-growth area. Metal-on-dielectric systems are more uncommon in the area-selective ALD literature compared to the selective dielectric-on-dielectric deposition.⁴⁸ Besides, there are very few reports on area-selective deposition of metal sulfides. Therefore, the process should be a valuable addition to the area-selective deposition toolbox.

Film growth on different substrates: toward epitaxial growth

In addition to the substrates described above, Co_9S_8 films were successfully grown on a variety of other dielectric, semiconducting, and metallic substrates, including sapphire ($\alpha\text{-Al}_2\text{O}_3$), mica, GaN, and different ALD-grown films (Al_2O_3 , Ir, and TiO_2). The morphologies of 1000 cycle films were similar on most of the substrates, although on sapphire and ALD- TiO_2 the size of the surface features was somewhat smaller and on ALD-Ir larger compared to the other substrates, *i.e.* native SiO_2 , thermal SiO_2 , Si-H, SLG, and ALD- Al_2O_3 (Fig. 4).

XRD revealed differences in the crystalline texture: the films on ALD- Al_2O_3 were more strongly (111) oriented compared to sapphire and native SiO_2 and even more so compared to Si-H, SLG, and thermal SiO_2 (ESI, Fig. S18†). On ALD-Ir and ALD- TiO_2 , the films appeared (113) textured, which may also be affected by the rougher morphology of these substrates.

Some differences in growth rate and resistivity between the different substrates studied were also observed (ESI, Fig. S19†). These may be due to differences in nucleation delay and the effect of crystallographic texture on growth rate (section S3 in the ESI†).

On single-crystalline mica and GaN substrates, a very different surface morphology mostly exposing triangular grains was observed. XRD showed the films to be very strongly (111) textured to a degree where no Co_9S_8 peaks were observed in the GIXRD measurements, and only strong (111), (222), and

(333) reflections were observed in the $\theta - 2\theta$ measurements. AFM confirmed the different morphology and larger size of the surface features on GaN and mica compared to Si (ESI, Fig. S20†).

The very strong (111) texture on mica and GaN may suggest an epitaxial alignment between Co_9S_8 and these single-crystalline substrates. First, it was confirmed that the out-of-plane peaks attributed to Co_9S_8 did not originate from the multi-layered GaN substrate comprised of graded Al(Ga)N buffer layers on Si(111) (ESI, Fig. S21a and b†). Then, the texture was quantified using a rocking curve (ω scan) measurement, which confirmed the strong (111) texture with a relatively small full-width at half-maximum (FWHM) of 1.5° (ESI, Fig. S21c†).

An in-plane pole figure measurement of the (111) Co_9S_8 reflection revealed a sharp pole in the middle, corresponding to the (111) out-of-plane orientation (Fig. 5a; extracted cut along $90^\circ\beta$ in Fig. S22a in the ESI†). In addition, six poles tilted by 70.5° with respect to the substrate plane (*i.e.* at $19.5^\circ\alpha$) were observed (see also Fig. S22b in the ESI†), which were attributed to the $\{\bar{1}11\}$ planes that are at an angle of 70.5° with respect to the (111) plane in the crystal structure of Co_9S_8 (cubic, space group 225, $Fm\bar{3}m$). The observation of six poles rather than a continuous ring confirms the epitaxial nature of the films and, as the $(\bar{1}11)$ reflection has only three-fold rotational symmetry which would result in three peaks every 120° , suggests a presence of two domains rotated by 60° with respect to each other.

A second pole figure of the substrate $(\bar{1}102)\text{GaN}$ reflection (Fig. 5b and Fig. S22c in the ESI†) revealed six sharp poles tilted by 43° with respect to the substrate ($47^\circ\alpha$) surface that were observed at the same β (sample rotation around its normal) angles as the $(\bar{1}11)\text{Co}_9\text{S}_8$ poles. Six poles are expected due to the six-fold rotational symmetry of wurtzite GaN (hexagonal, space group 186, $P6_3mc$, PDF 2-1078). The three sharp and weak poles observed at $35^\circ\alpha$ and six poles at $62^\circ\alpha$ were attributed to the substrate as confirmed by a measurement of a bare substrate but they could not be assigned to any known reflection. These extra poles may be due to multiple scattering effects or the non-monochromatic radiation used.

The epitaxial film-substrate registration was confirmed by in-plane measurements, which detect planes perpendicular to

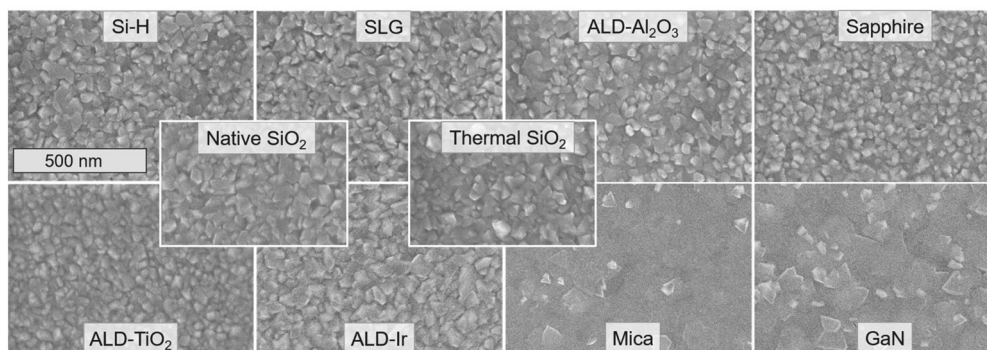


Fig. 4 SEM images of films deposited on different substrates at 275 °C using 1000 cycles.



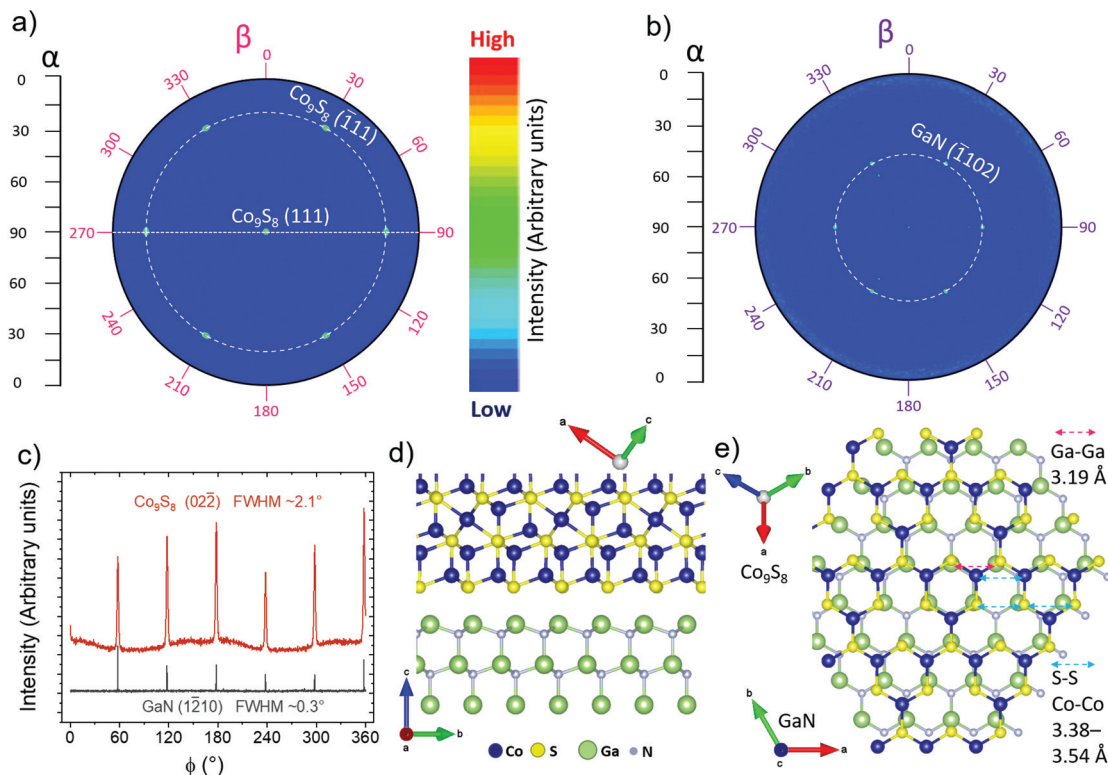


Fig. 5 In-plane pole figures at (a) $15.5^\circ 2\theta$ showing the out-of-plane $(111)\text{Co}_9\text{S}_8$ reflection in the middle ($90^\circ\alpha$) and $(\bar{1}11)$ reflections tilted by 70.5° ($19.5^\circ\alpha$) and (b) $48.27^\circ 2\theta$ showing the $(\bar{1}102)\text{GaN}$ reflections at $47^\circ\alpha$. Both pole figures were measured with the sample in the same orientation. (c) In-plane ϕ scans of $(022)\text{Co}_9\text{S}_8$ ($25.35^\circ 2\theta_\chi$) and $(\bar{1}210)\text{GaN}$ ($57.95^\circ 2\theta_\chi$) reflections. The data has been vertically offset for clarity. Schematic of a possible epitaxial arrangement $(111)[011]\text{Co}_9\text{S}_8||(\bar{0}001)[2\bar{1}\bar{1}0]\text{GaN}$: (d) cross-sectional and (e) top view. In top view, only one puckered Co–S “layer” is shown for clarity. Selected atomic distances are indicated.

the substrate surface. In the in-plane ϕ scans (full 360° rotation), six peaks from both $(022)\text{Co}_9\text{S}_8$ and $(\bar{1}210)\text{GaN}$ reflections were observed at the same ϕ angles (Fig. 5c). Both of these reflections have six-fold rotational symmetry, although any of the possible surfaces have, at most, three-fold symmetry. Whether one or two Co_9S_8 domains are present thus cannot be determined from this measurement. The presence of two domains was shown by the in-plane pole figure measurements discussed earlier, which is also in line with the observation of some approximately triangular grains rotated by 60° or 180° with respect to each other by SEM (Fig. 4). The in-plane pole figure and in-plane ϕ scan measurements therefore suggest epitaxial film-substrate relations $(111)[011]\text{Co}_9\text{S}_8||(\bar{0}001)[2\bar{1}\bar{1}0]\text{GaN}$ and $(111)[\bar{1}\bar{1}0]\text{Co}_9\text{S}_8||(\bar{0}001)[2\bar{1}\bar{1}0]\text{GaN}$, the two orientations being related by a 60° rotation.

The 2.1° FWHM of the $(022)\text{Co}_9\text{S}_8$ in-plane peaks is reasonably small, although clearly larger than that of the $(\bar{1}210)\text{GaN}$ reflection (0.3°), suggesting good epitaxial quality considering the low deposition temperature. Thinner films deposited on GaN using 100 and 250 cycles were also found to be epitaxial with similar in- and out-of-plane disorder to the 1000 cycle film (ESI, Fig. S23†). Furthermore, the 100 cycle film on GaN was already conductive at a thickness of approximately 5 nm. All of the epitaxial $\text{Co}_9\text{S}_8/\text{GaN}$ films were relatively rough with

roughnesses similar to the polycrystalline Co_9S_8 films deposited on silicon.

A look at the atomic structures of GaN and Co_9S_8 in the observed alignment shows obvious similarities that may explain the epitaxial growth between two materials which at the first sight have very different structures (a possible interface drawn in Fig. 5d, e and discussion in section S8 in the ESI†). An approximately 10% mismatch between the illustrated surfaces suggests that instead of conventional epitaxy, a domain matching epitaxy mechanism⁴⁹ may be in play (section S8 in the ESI†).

Of the other single-crystalline substrates evaluated, on mica epitaxial growth with several distinct domains and large in-plane disorder was observed, as discussed in the ESI (section S9).† Large disorder and multiple domains are rather common for van der Waals epitaxy, for which the film-substrate bonding is weak,^{50,51} suggesting that the $\text{Co}_9\text{S}_8/\text{mica}$ system may follow this mechanism. On sapphire, on the other hand, even the out-of-plane (111) texture of the Co_9S_8 films was not very strong, which rules out epitaxial alignment (section S10 in the ESI†).

We hope that the demonstration of the first epitaxial Co_9S_8 films inspires further studies on the epitaxial $\text{Co}_9\text{S}_8/\text{GaN}$ system including film structure, substrate/film interface, and



film properties. We believe that careful analysis of the system may both improve fundamental understanding of epitaxy of dissimilar materials as well as enable preparation of higher quality Co_9S_8 films with improved properties.

High-temperature stability in different atmospheres

High-temperature XRD (HTXRD) was used to evaluate the stability and reactivity of the Co_9S_8 films under different atmospheres. On one hand, high-temperature stability is required for various applications and device processing steps. On the other hand, post-deposition annealing can also be used to tune the composition of the films. Fig. 6 shows a summary of the phase transformations observed under inert, reducing, and oxidizing atmospheres as will be discussed below.

Under a flowing, reduced pressure (30 mbar) N_2 atmosphere, no changes in the crystallinity and phase composition were observed upon heating up to the highest evaluated temperature of 750 °C (ESI, Fig. S27†). The resistivity of the film decreased from 95 to 65 $\mu\Omega\text{ cm}$ and although minor changes in film morphology were observed, the film remained continuous after the annealing (ESI, Fig. S28†). In the published Co–S phase diagram, Co_9S_8 has been indicated to be stable up to 830 °C.³ The high stability, lack of reactions with silicon, and absence of strong agglomeration at these temperatures is encouraging considering possible use of Co_9S_8 in applications where conductive films are needed and high-temperature processing is encountered.

Under a dynamic vacuum (approximately 10^{-5} mbar), the Co_9S_8 film was stable up to 600 °C, where cubic CoSi (PDF 8-362) appeared to form (ESI, Fig. S29†). At 675 °C, fcc Co (PDF 15-806) and cubic CoSi_2 (PDF 38-1449) were also observed to form. After heating to 750 °C and cooling to room tempera-

ture, only fcc Co and CoSi_2 were observed (ESI, Fig. S30†). The formation of CoSi_x is likely preceded by loss of sulfur driven by the low pressure conditions and formation of either a sulfur-poor CoS_x phase or metallic Co, which can then diffuse through the native SiO_2 layer to react with the silicon substrate. Indeed, CoSi_2 has been reported to form by a reaction of Co thin film with silicon at 700 °C.^{52,53}

Under a reducing forming gas environment (10% H_2 /90% N_2 , atmospheric pressure), the Co_9S_8 films were stable up to approximately 400 °C, whereupon metallic cobalt started to form (Fig. 7). The film was completely transformed to cobalt at 475 °C, mainly in the fcc phase (PDF 15-806), with a minor contribution from hcp Co (PDF 5-727). Metallic Co remained until the end of the measurement at 725 °C. After cooling down to room temperature, the film consisted of mostly fcc Co with some hcp Co present (ESI, Fig. S31a†). In bulk, hcp Co is known as the low-temperature phase, which ideally transforms to the high-temperature fcc phase at 422 °C.³ The formed metallic film was discontinuous (ESI, Fig. S31b†), as is often observed after high-temperature reduction of compound films to metals.³² Interestingly, formation of CoSi_x was not observed, unlike in the measurement in vacuum. Notably, the temperature needed for reduction of Co_9S_8 was considerably higher compared to the reduction of CoO films grown from the same cobalt precursor (250 °C in the same atmosphere),³² which reflects the high stability of Co_9S_8 .

Finally, the stability of Co_9S_8 films under oxidizing conditions was evaluated by annealing in air (Fig. 8). No changes were observed up to 225 °C, whereupon Co_9S_8 started to convert to Co_3S_4 (linnaeite, PDF 42-1448). Formation of Co_3O_4 (PDF 42-1467) was observed at 275 °C, and the oxide coexisted with various sulfide and sulfate phases during the measurement: with Co_3S_4 up to 300 °C, with Co_3S_4 and CoS_2 (cattierite, PDF 41-1471) at 325–375 °C, and with CoSO_4 (PDF 46-4) at 400–550 °C until Co_3O_4 was the only phase present at 575 to 725 °C.

One potential explanation for the formation of phases with a higher sulfur content compared to Co_9S_8 , *i.e.* Co_3S_4 and CoS_2 , may be release of elemental sulfur upon gradual oxidation of Co_9S_8 to Co_3O_4 . While sulfur will eventually oxidize to SO_x under oxidizing conditions, we hypothesize that at low enough temperatures, part of the elemental sulfur can survive long enough to form more sulfur-rich CoS_x phases by reacting with Co_9S_8 . At higher temperatures in oxygen-rich atmosphere, these sulfide phases may be further oxidized to CoSO_4 , while the excess sulfur can be released as gaseous SO_x . Similarly, decomposition of CoSO_4 to Co_3O_4 may release the sulfur as SO_x species.

After cooling down to room temperature, a rough but continuous Co_3O_4 film was obtained (ESI, Fig. S32†). The observed oxidation and phase transition temperatures are similar to a previous study on Co_3S_4 thin films made by chemical bath deposition.⁵⁴

Under an atmospheric pressure of pure oxygen, the high-temperature reactions were similar to those in laboratory air, although some of the phase transitions occurred at slightly

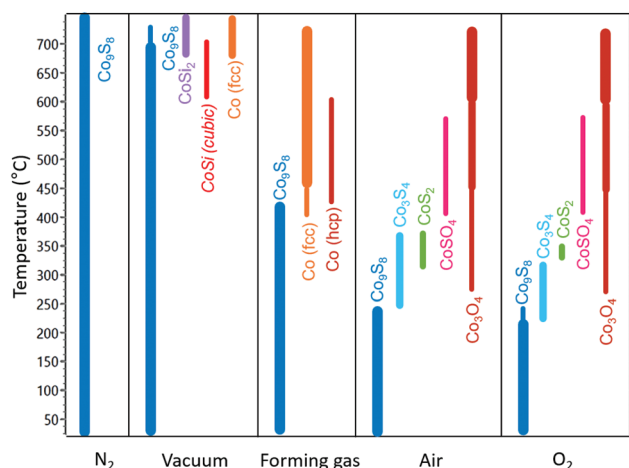


Fig. 6 Summary of crystalline phases observed during *in situ* high-temperature GIXRD measurements performed under different atmospheres. The widths of the lines referring to the different phases represent an estimation on the relative amounts of the different phases based on peak intensities.



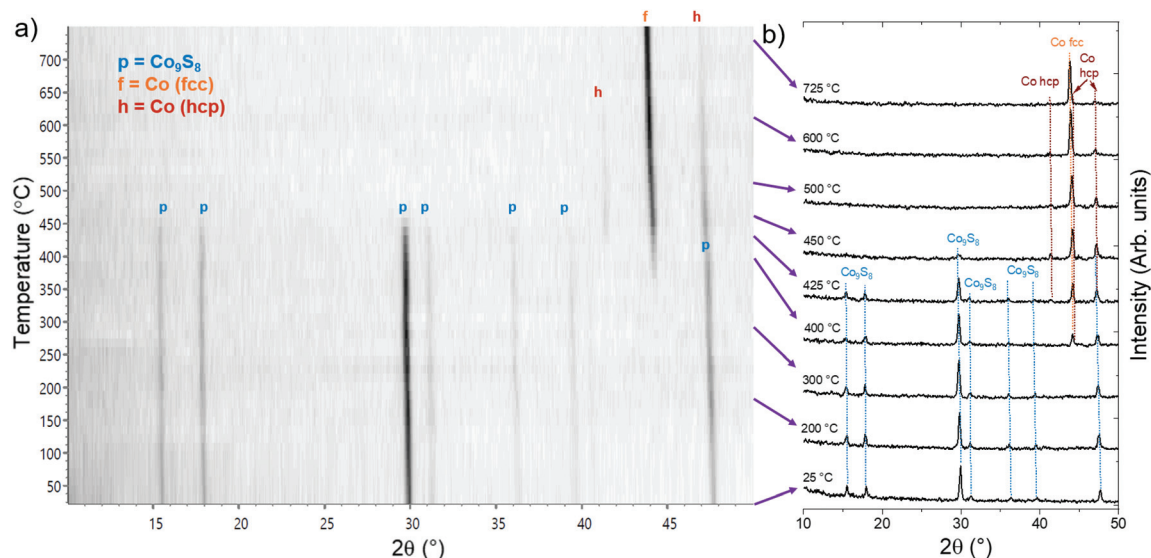


Fig. 7 (a) High-temperature grazing incidence X-ray diffractograms during heating in a forming gas atmosphere (10% H_2 /90% N_2 , atmospheric pressure) from 25 to 725 °C (measured at 25 °C intervals), where light and dark represent low and high intensities, respectively. (b) Selected diffractograms measured at different temperatures. Peaks attributed to different phases were identified by symbols and lines of different colors. The studied film was deposited on silicon (native oxide) at 275 °C using 1000 cycles.

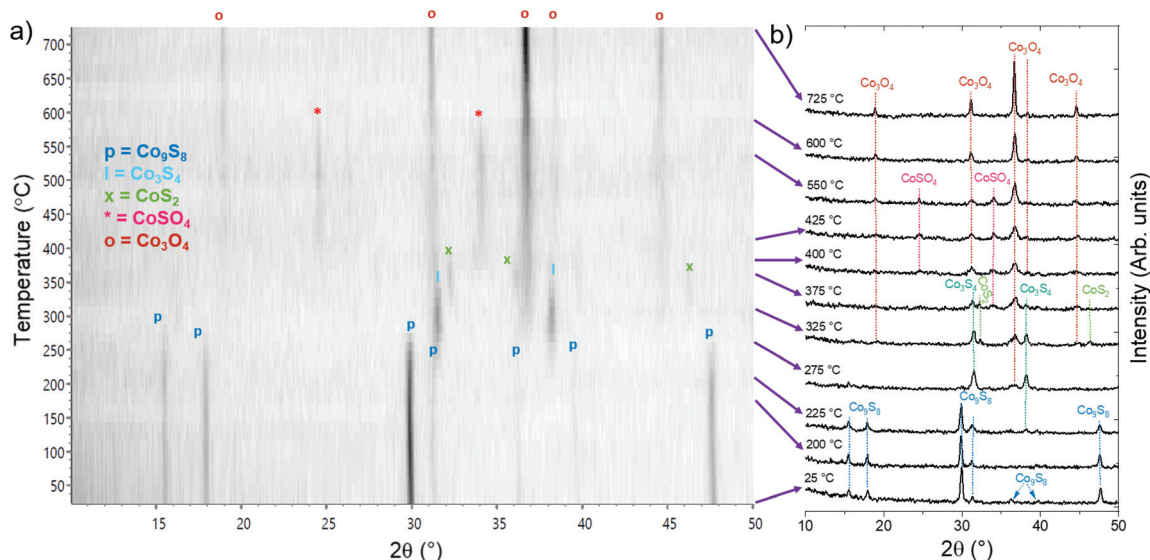


Fig. 8 (a) High-temperature grazing incidence X-ray diffractograms during heating in air (atmospheric pressure) from 25 to 725 °C (measured at 25 °C intervals), where light and dark represent low and high intensities, respectively. (b) Selected diffractograms measured at different temperatures. Peaks attributed to different phases were identified by symbols and lines of different colors. The studied film was deposited on silicon (native oxide) at 275 °C using 1000 cycles.

lower temperatures (ESI, Fig. S33†) and the resulting oxide film appeared to be more strongly crystallized (ESI, Fig. S34†).

Conclusions

In this work, we have developed a new ALD process for Co_9S_8 thin films using $\text{CoCl}_2(\text{TMEDA})$ with H_2S . Compared to the

earlier CoS_x ALD processes, the precursor $\text{CoCl}_2(\text{TMEDA})$ is more easily synthesized using inexpensive starting materials, and the resulting Co_9S_8 films have a considerably lower resistivity. Using our new process, highly conductive, pure, and crystalline Co_9S_8 films can be deposited between 180 and 300 °C. The lowest resistivity (80 $\mu\Omega\text{ cm}$), best uniformity, and highest growth rate are achieved at 275 °C. Electrically conductive films as thin as approximately 3 nm thick can be prepared



on silicon with a native oxide layer. Area-selective deposition is enabled by fast nucleation on oxide-covered silicon in contrast to at least 50 cycle nucleation delay on Si-H, enabling deposition of continuous, at least 5 nm thick films without any growth on Si-H. Besides silicon oxide, Co₉S₈ films can also be deposited on different oxide, metal, and nitride substrates. The first ever epitaxial Co₉S₈ films were obtained on GaN, as studied in detail by XRD. The Co₉S₈ films are stable up to at least 750 °C in N₂, 600 °C in high vacuum, 400 °C in forming gas, and 225 °C in air or pure O₂. The high stability is promising for use in applications requiring high-temperature processing steps. Overall, we believe the presented process may be interesting for, for example, metal barrier, liner, and seed layers for interconnects as well as various energy technology applications due to the high conductivity of the films and scalable, simple, low cost synthesis of the cobalt precursor.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

The research was supported by ASM Microchemistry. Imec is thanked for supplying the epitaxial GaN substrates.

Notes and references

- C. N. R. Rao and K. P. R. Pisharody, *Prog. Solid State Chem.*, 1976, **10**, 207.
- Y. O. Chen and Y. A. Chang, *Metall. Trans. B*, 1978, **9**, 61.
- H. Rau, *J. Phys. Chem. Solids*, 1976, **37**, 931.
- D. Xiong, Q. Zhang, S. M. Thalluri, J. Xu, W. Li, X. Fu and L. Liu, *Chem. – Eur. J.*, 2017, **23**, 8749.
- X. Ma, W. Zhang, Y. Deng, C. Zhong, W. Hu and X. Han, *Nanoscale*, 2018, **10**, 4816.
- H. Li, Z. Guo and X. Wang, *J. Mater. Chem. A*, 2017, **5**, 21353.
- L.-L. Feng, M. Fan, Y. Wu, Y. Liu, G.-D. Li, H. Chen, W. Chen, D. Wang and X. Zou, *J. Mater. Chem. A*, 2016, **4**, 6860.
- D. Kim, J.-G. Song, H. Yang, H. Lee, J. Park and H. Kim, *Nanoscale*, 2019, **11**, 844.
- Y. Chen, J.-M. Tarascon and C. Guéry, *Electrochim. Acta*, 2013, **99**, 46.
- L. Zheng, X. Sun, L. Chen, C. Bao, W. Luo, N. Huang, P. Sun, Y. Sun, L. Fang and L. Wang, *J. Mater. Sci.*, 2016, **51**, 4150.
- S.-H. Chang, M.-D. Lu, Y.-L. Tung and H.-Y. Tuan, *ACS Nano*, 2013, **7**, 9443.
- S. Y. Chae, Y. J. Hwang, J.-H. Choi and O.-S. Joo, *Electrochim. Acta*, 2013, **114**, 745.
- H. Li, Y. Gao, Y. Shao, Y. Su and X. Wang, *Nano Lett.*, 2015, **15**, 6689.
- W. Xiong, K. Hu, Z. Li, Y. Jiang, Z. Li, Z. Li and X. Wang, *Nano Energy*, 2019, **66**, 104149.
- C. Auth, *et al.*, in *2017 IEEE International Electron Devices Meeting (IEDM)*, 2017, p. 673.
- X. F. Qian, X. M. Zhang, C. Wang, Y. Xie and Y. T. Qian, *Inorg. Chem.*, 1999, **38**, 2621.
- N. Kumar, N. Raman and A. Sundaresan, *Z. Anorg. Allg. Chem.*, 2014, **640**, 1069.
- Y. N. Ko, S. H. Choi, S. B. Park and Y. C. Kang, *Chem. – Asian J.*, 2014, **9**, 572.
- K. Ramasamy, W. Maneeprakorn, N. Iqbal, M. A. Malik and P. O'Brien, *Int. J. Nanosci.*, 2011, **10**, 815.
- V. G. Bessergenev, R. J. F. Pereira and A. M. Botelho do Rego, *Surf. Coat. Technol.*, 2007, **201**, 9141.
- R. Nomura and N. Nakai, *Trans. Mater. Res. Soc. Jpn.*, 2003, **28**, 1287.
- S. M. George, *Chem. Rev.*, 2010, **110**, 111.
- R. W. Johnson, A. Hultqvist and S. F. Bent, *Mater. Today*, 2014, **17**, 236.
- M. Ritala and J. Niinistö, in *Chemical Vapor Deposition: Precursors, Processes and Applications*, ed. A. C. Jones and M. L. Hitchman, Royal Society of Chemistry, Cambridge, 2009, pp. 158–206.
- M. Ritala and J. Niinistö, *ECS Trans.*, 2009, **25**(8), 641.
- I. J. Raaijmakers, *ECS Trans.*, 2011, **41**(2), 3.
- Atomic Limits ALD Database, <https://www.atomiclimits.com/alddbatabase/> (accessed August 2021).
- N. P. Dasgupta, X. Meng, J. W. Elam and A. B. F. Martinson, *Acc. Chem. Res.*, 2015, **48**, 341.
- A. W. Peters, Z. Li, O. K. Farha and J. T. Hupp, *ACS Nano*, 2015, **9**, 8484.
- Z. Guo and X. Wang, *Angew. Chem., Int. Ed.*, 2018, **57**, 5898.
- Strem Chemicals, Inc., Product Catalog: Bis(*N,N'*-di-*i*-propylacetamidinato)cobalt(II), https://www.strem.com/catalog/v/27-0486/16/cobalt_635680-58-9 (accessed June 2021).
- K. Väyrynen, T. Hatanpää, M. Mattinen, M. Heikkilä, K. Mizohata, K. Meinander, M. Ritala and M. Leskelä, *Chem. Mater.*, 2018, **30**, 3499.
- K. Väyrynen, T. Hatanpää, M. Mattinen, K. Mizohata, K. Meinander, J. Räisänen, J. Link, R. Stern, M. Ritala and M. Leskelä, *Adv. Mater. Interfaces*, 2018, **6**, 1801291.
- H.-E. Nieminen, M. Kaipio and M. Ritala, *Chem. Mater.*, 2020, **32**, 8120.
- T. Suntola, *Thin Solid Films*, 1992, **216**, 84.
- N. P. Dasgupta, J. F. Mack, M. C. Langston, A. Bousetta and F. B. Prinz, *Rev. Sci. Instrum.*, 2010, **81**, 044102.
- W. de Poel, S. Pintea, J. Drnec, F. Carla, R. Felici, P. Mulder, J. A. A. W. Elemans, W. J. P. van Enckevort, A. E. Rowan and E. Vlieg, *Surf. Sci.*, 2014, **619**, 19.
- P. R. Ribic and G. Bratina, *Surf. Sci.*, 2007, **601**, 44.
- R. A. Waldo, in *Microbeam Analysis*, ed. D. E. Newbury, San Francisco Press, San Francisco, 1988, pp. 310–314.
- Co₉S₈ Crystal Structure, SpringerMaterials (online database), ed. P. Villars and K. Cenzual, Springer, Heidelberg, 2016, https://materials.springer.com/isp/crystallographic/docs/sd_0375057 (accessed June 2021).



- 41 K. Nagao and E. Kagami, *Rigaku J.*, 2011, **27**, 6.
- 42 S. Kobayashi, *Rigaku J.*, 2010, **26**, 3.
- 43 K. Momma and F. Izumi, *J. Appl. Crystallogr.*, 2011, **44**, 1272.
- 44 *CRC Handbook of Chemistry and Physics*, ed. J. R. Rumble, CRC Press, Boca Raton, FL, 102nd edn, 2021, <http://www.hbcpnetbase.com/> (accessed June 2021).
- 45 K. Väyrynen, PhD thesis, University of Helsinki, 2019. Available at <https://helda.helsinki.fi/handle/10138/305438>.
- 46 L. T. Zhuravlev, *Langmuir*, 1987, **3**, 316.
- 47 R. B. Milligan and D. Li, *ECS Trans.*, 2010, **33**(2), 159.
- 48 G. N. Parsons and R. D. Clark, *Chem. Mater.*, 2020, **32**, 4920.
- 49 J. Narayan and B. C. Larson, *J. Appl. Phys.*, 2003, **93**, 278.
- 50 M. Mattinen, P. J. King, G. Popov, J. Hämäläinen, M. J. Heikkilä, M. Leskelä and M. Ritala, *2D Mater.*, 2020, **7**, 011003.
- 51 M. Nakano, Y. Wang, Y. Kashiwabara, H. Matsuoka and Y. Iwasa, *Nano Lett.*, 2017, **17**, 5595.
- 52 J.-H. Yoon, H.-B.-R. Lee, G.-H. Gu, C.-G. Park and H. Kim, *Korean J. Mater. Res.*, 2012, **22**, 202.
- 53 C. T. Joensson, I. A. Maximov, H. J. Whitlow, V. Shutthanandan, L. Saraf, D. E. McCready, B. W. Arey, Y. Zhang and S. Thevuthasan, *Nucl. Instrum. Methods Phys. Res., Sect. B*, 2006, **249**, 532.
- 54 F. C. Eze, *Semicond. Sci. Technol.*, 2001, **16**, 362–366.

