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The origin of the unusual red-shifted aggregation-state emission of triphenylamine-imidazole molecules: excimers or a photochemical reaction?†

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At the moment, there is still no consensus regarding the origin of the newly apparent unusual red-shifted aggregation-state emission arising from some triphenylamine-imidazole luminescent materials. One view is that the unusual emission originates from the formation of excimers, but a recently observed photochemical reaction also seems to contribute to it. Herein, a new imidazole molecule, **TPDBM**, was synthesized to investigate the actual origin of this phenomenon. Compared with a dilute solution of **TPDBM**, its film obviously exhibits newly apparent high red-shifted emission at around 520 nm in the PL spectrum. According to UV-vis spectra and aggregation-induced emission (AIE) and concentration-dependent emission results, it seems that the red-shifted emission is derived from excimers. However, further studies of excitation spectra, *in situ* PL spectra and oxygen-free PL spectra entirely overturn this conclusion and ultimately confirm that it seems to come from a photochemical reaction occurring on the imidazole ring of **TPDBM**. Two photochemical products, **TPBMZ** and **TPDYZ**, with excellent AIE effects were obtained through this photochemical reaction, and **TPBMZ** can be further oxidized to **TPDYZ** due to the quite fast reaction rate. Furthermore, the photochemical reaction of **TPDBM** doped into PMMA film was suitably applied as a fluorescent probe for oxygen with a lowest detectable oxygen volume ratio of 0.1%.

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Introduction

As a classical molecular structural unit of functional materials,^{1–3} imidazole possesses a typical five-member ring containing two entirely different nitrogen atoms, where one is electron-rich and the other is electron-deficient, which allows great control of the photoelectrical properties of imidazole derivatives. This is why imidazole derivatives can be extensively applied to OLEDs,^{4–12} fluorescent switches,^{13–15} optical sensors,^{16–18} electron transfer mediators,¹⁹ memory devices²⁰ and other fields of interest.^{21–26} With the requirement for improved material performance, it is even more necessary to further explore imidazole molecules with new structures and investigate their photophysical and photochemical properties.

Triphenylamine, as a well-known electron donor,^{27–32} has been introduced to build imidazole molecules^{33–36} that have shown excellent photoelectrical properties, such as high quantum efficiencies,³⁵ balanced charge mobilities,³⁶ and so on. Many interesting and impressive studies have been already reported describing the luminescence properties of triphenylamine-imidazole molecules.^{37–41} For example, Li *et al.* developed a molecule with deep-blue emission, **TPA-PIM**, with an extraordinarily narrow FWHM of 35 nm.³⁸ And a similar molecule (**PTN**) was reported by Du *et al.* with newly apparent unusual red-shifted emission in the aggregation state, which was attributed to excimer formation by the authors.³⁹ Recently, a photochemical reaction on the imidazole ring of this kind of molecule (**DPA-PIM**) was reported,⁴⁰ and the photochemical products also displayed high red-shifted emission in the aggregation state with yellow fluorescence under UV and air conditions. Based on these studies, it is necessary to undertake deep and careful investigations to further confirm whether this kind of unusual emission from imidazole molecules is due to excimers or a photochemical reaction.

Herein, we designed and successfully synthesized a new triphenylamine-imidazole molecule, **TPDBM**, with two imidazole rings linked on one triphenylamine unit. An obvious large red-shifted

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emission peak at around 520 nm appeared in the PL spectrum of **TPDBM** film compared to in solution. We discussed the source of this unusual emission based on spectroscopic analysis and investigations of photochemical reactions involving **TPDBM**, and finally arrived at the conclusion that this unusual red-shifted emission comes from a photochemical reaction occurring on the imidazole ring of **TPDBM** and not from excimer formation. Furthermore, two photochemical products, **TPBMZ** and **TPDYZ**, were obtained through this photochemical reaction, and **TPBMZ** could be further turned to **TPDYZ** under the same conditions due to the quite fast reaction rate. Surprisingly, both of the photochemical products exhibited excellent aggregation-induced emission (AIE) properties. Moreover, the photochemical reaction involving **TPDBM** doped into PMMA film was applied to a fluorescent probe for oxygen, which displayed excellent performance with a lowest detectable oxygen volume ratio of 0.1%.

Results and discussion

Material synthesis and characterization

As shown in Fig. 1, the triphenylamine-imidazole molecule **TPDBM** was prepared simply *via* a cyclization reaction⁴⁰ from benzil, aniline, ammonium acetate and the intermediate compound 4,4'-(phenylazanediyl)dibenzaldehyde (**M1** in Fig. 1), which was newly synthesized *via* an improved process with a yield of 80% according to the related ref. 42. The structure of the target product **TPDBM** was confirmed *via* ¹H NMR and mass spectroscopy studies. All the reaction details relating to this work can be found in the Experimental section.

Photophysical properties

UV-vis and PL spectroscopy studies were carried out to investigate the basic photophysical properties of **TPDBM**. As shown in Fig. 2, when **TPDBM** was excited, significantly different PL spectra were obtained from **TPDBM** in solution and as a film. A diluted toluene solution of **TPDBM** (10^{-5} M) exhibited a relatively narrow single emission peak at around 405 nm. But as a film, except for the normal emission peak at around 430 nm, another large red-shifted emission peak at around 520 nm was also observed. Furthermore, when the above-mentioned **TPDBM** film possessing dual-emission properties was re-dissolved in toluene, the emission at around 520 nm disappeared and the spectrum became identical to the previous regular toluene solution. This transformation process can be well observed from the fluorescence images shown in the inset of Fig. 2. Also, the UV-vis spectra of **TPDBM** both in toluene and as



Fig. 2 UV-vis and PL spectra of **TPDBM** in toluene solution and film states; the inset shows the excitation spectra of **TPDBM** film at 430 nm and 520 nm.

a film exhibited almost identical broad absorption bands located at around 365 nm, indicating that the emission of **TPDBM** seems to originate from the same ground state whether in a solution or film state. Thus, this unusual large red-shifted emission peak of **TPDBM** in the film state seems to be related to the molecular aggregation state.

To further investigate the influence of the degree of molecular aggregation on the formation of this large red-shifted emission peak, PL spectra of **TPDBM** in THF and H₂O mixed solutions with different water fraction (f_w) levels were measured (Fig. 3a). As **TPDBM** is almost insoluble in H₂O, it aggregates as the amount of H₂O increases; therefore, the emission behaviour resulting from different aggregation degrees can provide clues for investigating the luminescence properties of **TPDBM**. **TPDBM** in dilute THF solution exhibited strong deep-blue emission at around 405 nm and its PL intensity remained nearly unchanged in the THF/H₂O mixture solution at a low water fraction. As the water fraction increased to 80%, the emission peak at around 405 nm started to become weaker, whereas a new emission peak appeared at around 525 nm and further intensified as the water fraction was increased to 90%. Meanwhile the fluorescence of the mixed solution changed from blue to green. This result indicated that a large degree of molecular aggregation could lead to the formation of red-shifted emission. Moreover, PL spectra of **TPDBM** in chloroform solutions of different molarity were also obtained (Fig. 3b). All the PL spectrum intensities were normalized for the convenient analysing of differences between the emission peaks. When **TPDBM** existed at a lower molarity ($\leq 10^{-4}$ M), the PL spectrum exhibited only a single emission peak at around 412 nm. As the molarity was increased to 5×10^{-4} M and 10^{-3} M, a new red-shifted emission peak appeared at around 530 nm, which could also be described as resulting from an increase in the degree of molecular aggregation. So far, all the above-mentioned results demonstrated that the unusual large red-shifted emission peak was indeed related to a large degree of **TPDBM** molecular aggregation, which seems to be attributed to

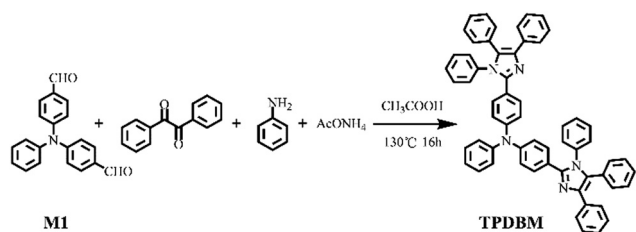


Fig. 1 The molecular structure and synthesis route of **TPDBM**.



Fig. 3 (a) PL spectra of **TPDBM** in mixed solutions of THF and H₂O; the inset exhibits the actual luminescence of solutions with different water fractions. (b) PL spectra of **TPDBM** in chloroform solutions of different molarity. (c) *In situ* PL spectra of **TPDBM** in dilute toluene (10^{-5} M) solutions under UV light (365 nm) for different periods of photo-irradiation; the emission wavelength spectra from 520 to 680 nm are magnified in the inset. (d) *In situ* PL spectra of **TPDBM** in film form.

the formation of excimers in the aggregated state and consistent with the report by Du *et al.*

However, as we looked at the excitation spectra of **TPDBM** films at 430 and 520 nm shown in Fig. 2, we were surprised to find that there was a marked difference between them, *i.e.*, the former has only one peak at around 390 nm, and the latter has two peaks at around 365 and 415 nm. Unlike the UV spectra of **TPDBM**, this result indicates that the two emission peaks of **TPDBM** film come from different ground states, and the unusual large red-shifted emission peak from **TPDBM** film should not simply be described as arising from excimer emission. Further studies of *in situ* PL spectra of **TPDBM** in diluted toluene solution (10^{-5} M) also offered us new insights. As shown in Fig. 3c, as the photo-irradiation time increases, the PL intensity of the emission peak at around 405 nm gradually decreases; a small emission peak also appears at around 560 nm and becomes gradually enhanced. Since it is impossible to form excimers in dilute solutions, we could deduce that the red-shifted emission here could not arise from excimer emission but is more likely the result of photoinduced chemical changes following UV-irradiation.⁴⁰ In addition, similar results were also obtained from *in situ* PL spectra of **TPDBM** in film form (Fig. 3d). As the irradiation time increased, the emission peak at around 420 nm kept getting weaker and almost completely disappeared after 180 s. At the same time, the emission peak at around 520 nm was also gradually weakened at a

relatively slow speed. These spectral changes indicated that **TPDBM** with strong deep-blue emission was constantly consumed during UV-irradiation and a new substance with relatively weak yellow emission was produced. These results clearly demonstrated that large levels of molecular aggregation are not the only necessary condition for the large red-shifted emission peak, and that a photochemical reaction involving **TPDBM** under UV-irradiation could also lead to this emission.

In order to further determine whether the actual origin of this large red-shifted emission peak of **TPDBM** in the film state relates to excimers or a photochemical reaction, PL spectra of **TPDBM** film in a pure N₂ environment were measured. As shown in Fig. 4a, totally different to in air, the PL spectrum of **TPDBM** film in pure N₂ exhibited only one emission peak at around 430 nm. Moreover, when this film was returned to air, the large red-shifted emission peak at around 520 nm appeared again and the resulting PL spectrum exhibited identical dual-emission to that measured originally in air. These results indicated that the presence of oxygen replaced the aggregation state as the indispensable condition for the generation of the unusual emission from the film; that is, the actual origin of this red-shifted emission is not likely to be excimer formation but is more likely to be the occurrence of a photochemical reaction. To further prove this point, we obtained three groups of *in situ* UV-vis spectra from **TPDBM** toluene solution (Fig. 4b). When the sample

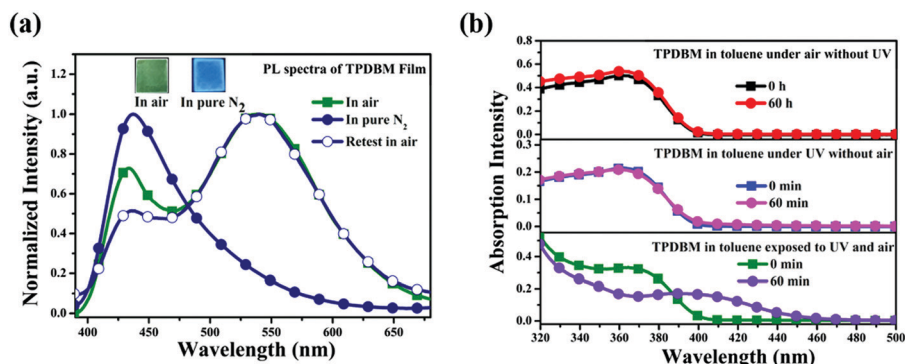


Fig. 4 (a) PL spectra of TPDBM film in air and in a pure N_2 environment; the inset shows fluorescence images of TPDBM film in air and pure N_2 . (b) *In situ* UV spectra of TPDBM in toluene.

was kept in darkness without UV-irradiation or sealed in N_2 without exposure to O_2 , the UV-vis spectra remained unchanged. But when it was simultaneously exposed to UV and O_2 for 60 min, the absorption changed obviously with a new peak appearing at around 410 nm, which was consistent with the red-shifted emission seen in the excitation spectra. These results further supported the deduction that a photochemical reaction induced the red-shifted emission and pointed out that the two essential conditions for this photochemical reaction are the presence of UV-irradiation and oxygen from air.

Photochemical reaction

To identify how a photochemical reaction induced this unusual red-shifted emission, UV-irradiation experiments involving

TPDBM were carried out to obtain more information about this photochemical reaction and its products, as shown in Fig. 5. A high concentration of TPDBM dissolved in toluene could directly produce (*Z*)-*N*-((benzoylimino)(4-(phenyl(4-(1,4,5-triphenyl-1*H*-imidazol-2-yl)phenyl)amino)phenyl)methyl)-*N*-phenylbenzamide (TPBMZ) in 40% yield and *N,N'*-((1*Z*,1'*Z*)-((phenylazanediy)bis(4,1-phenylene))-bis((benzoylimino)methylene))bis(*N*-phenylbenzamide) (TPDYZ) in 20% yield under 365 nm UV-irradiation and in air (Fig. 5a); the molecular structures were confirmed *via* 1H NMR and mass spectroscopy studies. A single crystal of TPBMZ (CCDC number: 1969967 \dagger) was obtained *via* the slow diffusion of *n*-hexane solvent into chloroform solution (Fig. 5b); this clearly shows the configuration, with one imidazole of TPDBM undergoing ring-opening, and also verifies that a photochemical reaction indeed occurred

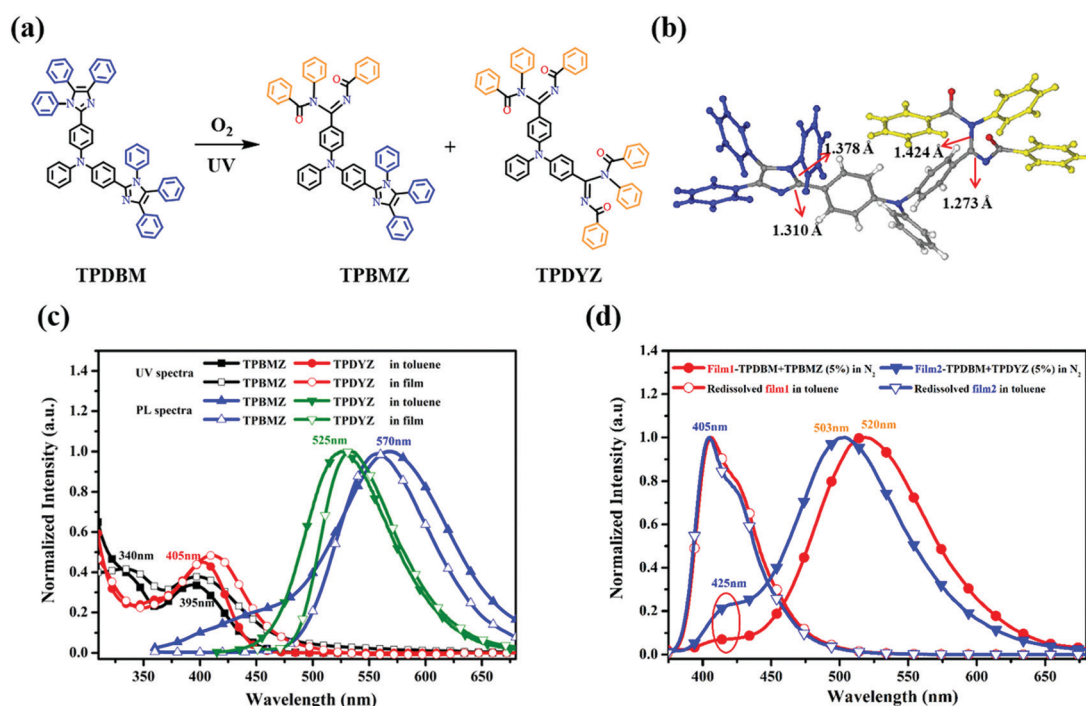


Fig. 5 (a) The photochemical reaction changing the molecular structure from TPDBM to TPBMZ and TPDYZ. (b) The single crystal structure of TPBMZ. (c) The UV and PL spectra of TPBMZ and TPDYZ both in toluene solution and film states. (d) The PL spectra of TPDBM doped with 5% (molarity) TPBMZ and TPDYZ in film form under pure N_2 and in toluene solution under air.

on the imidazole ring of **TPDBM**. Additionally, the structure of the **TPBMZ** crystal shows bond lengths of 1.27 Å for C=N and 1.42 Å for C-N, effectively confirming that benzoylimino and benzamide structures were generated *via* the ring-opening reaction.⁴⁰

The basic photophysical properties of the two photochemical products were also studied using their UV and PL spectra (Fig. 5c). In the UV spectrum of **TPBMZ** dissolved in toluene, the absorption peak of **TPDBM** at around 365 nm blue-shifted to 340 nm, while a new wide absorption band at around 395 nm, corresponding to the ring-opening structure of the imidazole ring, was also observed. On account of the disappearance of the imidazole-ring structure, the other product, **TPDYZ**, showed only one absorption peak at around 405 nm. Interestingly, the PL spectrum of **TPBMZ** in toluene exhibited dual emission, with two peaks at around 455 and 570 nm, respectively, which was probably caused by the special molecular structure, while **TPDYZ** showed a single emission peak at around 525 nm. The UV spectra of the two photochemical products in film form were very similar to those in toluene, with a slight red shift, and their PL spectra were virtually unchanged, except that the emission peak at around 455 nm from **TPBMZ** disappeared in the film state. The resulting UV and PL spectra from these two photochemical products were in very good agreement with the observed red-shifted emission and excitation spectra of **TPDBM** film.

Since the unusual red-shifted emission was indeed induced by a photochemical reaction involving the imidazole unit, why could it be observed obviously in the film and aggregation states but not in regular dilute solution? As shown in Fig. 5d, we prepared two types of **TPDBM** film doped with 5% (molarity) **TPBMZ** and **TPDYZ**, respectively, and tested their luminescence behaviour. Similar to pure **TPDBM** film in air, and regardless of whether the **TPDBM** film was doped with **TPBMZ** or **TPDYZ**, their PL spectra under pure N₂ both exhibited weak blue emission at around 425 nm and strong red-shifted emission located at around 520 or 503 nm, respectively. This should be ascribed to efficient Förster resonance energy transfer in the films due to the high levels of overlap between the PL spectrum of **TPDBM** (405 nm, Fig. 2) and the absorption spectra of the two photochemical products (395 and 405 nm, Fig. 5c). Moreover, when these two films were re-dissolved in toluene in air, both of them exhibited only one emission peak at around 405 nm, indicating that the impact of trace amounts of **TPBMZ** and **TPDYZ** dopants on the PL spectra of **TPDBM** in dilute solution is negligible. These results indicated that the unusual red-shifted emission of **TPDBM** in the aggregation state was actually emission from a small number of photochemical products excited by a Förster resonance energy transfer process, which is more effective in the aggregation state than in dilute solution. Therefore, although irradiation from the excitation light source unavoidably produced small amounts of photochemical products both in film and in toluene solution during the collection of PL spectra in air, the emission from **TPDBM** only exhibited obviously red-shifted emission peaks in film and not in toluene. This is why the emission peaks from **TPDBM** were so different between the toluene solution and film states, as shown in Fig. 3c and d.

To further explore this photochemical reaction, *in situ* UV spectra of **TPDBM** in toluene were obtained to monitor the specific reaction process. As shown in Fig. 6a, as the photo-irradiation time increased, the absorption peak at around 365 nm was gradually weakened and blue-shifted to 340 nm. Meanwhile, a new wide absorption peak at around 410 nm appeared and gradually increased in size. In addition, *in situ* UV spectra of **TPDBM** in the film state also exhibited a similar trend with a relatively faster reaction speed (Fig. 6b). These results clearly demonstrated that **TPDBM** gradually turned into the photochemical products **TPBMZ** and **TPDYZ** during this UV-irradiation process. The most noteworthy point is that there were no other special peaks apart from those of **TPDBM** and its photochemical products during this process, whether in toluene solution or film form, indicating that no by-products are generated. In addition, *in situ* ¹H NMR spectra of **TPBMZ** were also obtained to further describe this photochemical reaction process (Fig. 6c). The characteristic chemical shifts of pure **TPBMZ** and **TPDYZ** were distinctly demonstrated using their original ¹H NMR spectra in toluene-d₈ solution. The ¹H NMR spectrum of **TPBMZ** was measured again after being exposed to UV light in air for 90 min and the peak intensities at 8.07 and 7.58 ppm were obviously decreased, which clearly showed that **TPBMZ** turned to **TPDYZ** during the UV-irradiation process. This powerfully demonstrated that **TPBMZ** as a photochemical reaction product could further turn to the other product **TPDYZ** under the same conditions. In addition, as we know, the photochemical reaction occurred on the imidazole ring structure. Further investigations into the natural transition orbital (NTO) distributions of **TPBMZ** showed that a suitable structure was located in the high-energy S₃ excited state of **TPBMZ** (Fig. 6d). According to Kasha's rule, we inferred that the rate of this reaction is really very fast because it could occur even in the presence of very fast intersystem crossing processes from high-energy excited states to low-energy excited states.

AIE phenomenon

The AIE properties of the two photochemical products **TPBMZ** and **TPDYZ** were also investigated. Surprisingly, unlike **TPDBM**, both of the photochemical products exhibited obvious AIE phenomena, as revealed by the data and the photographs shown in Fig. 7. The PL spectrum of **TPBMZ** in pure THF showed weak dual emission with two peaks at 445 nm and 565 nm (Fig. 7a). In the THF/H₂O mixed solution system, as the water fraction increased, the PL peak intensity at 565 nm was greatly enhanced, accompanied by a slight blue shift to 560 nm. Meanwhile, the PL peak intensity at 445 nm continuously decreased and even almost disappeared, which further indicated that this weak emission perhaps comes from the imidazole-ring structure without AIE activity located in the higher energy excited state. Similar AIE behaviour was also observed from **TPDYZ** (Fig. 7b), which further demonstrated that the AIE emission from the photochemical products could be attributed to the generation of benzoylimino and benzamide groups as a result of ring opening. In addition, the AIE properties of these two photochemical products further explained why there were no visible red-shifted emission peaks in the PL spectra of **TPDBM** in dilute toluene solution, despite the presence of small amounts of the photochemical products.

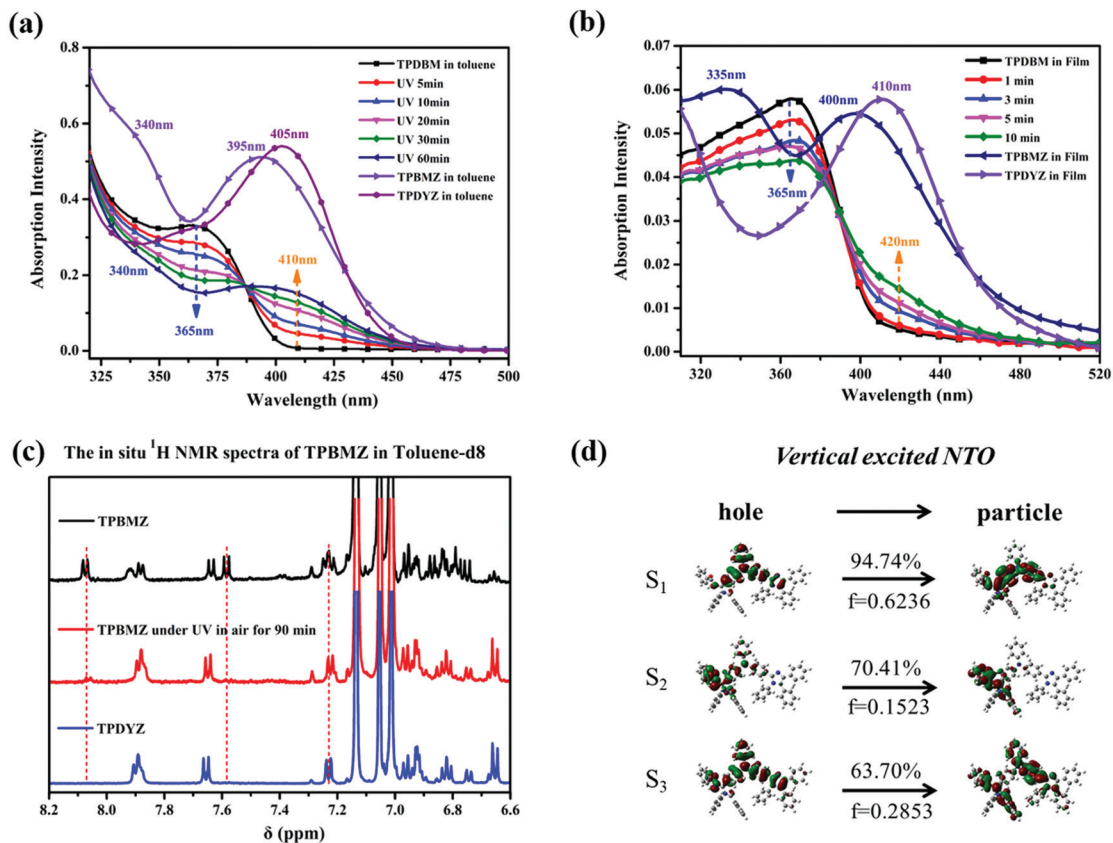


Fig. 6 *In situ* UV spectra of TPDBM (a) in toluene solution and (b) as a film. (c) *In situ* ¹H NMR spectra of TPBMZ. (d) The natural transition orbital (NTO) distributions of TPBMZ relating to S₀ → S₁, S₀ → S₂ and S₀ → S₃ excitations at the TD-ob97X/6-31G(d,p) level using S₀ state geometry.

Oxygen detection

Considering the outstanding sensitivity of TPDBM to oxygen and the quite rapid photochemical reaction rate, it could easily be applied to act as a fluorescent probe for oxygen. Furthermore, in order to meet the operational and low toxicity requirements for oxygen-sensing materials for practical applications, we doped TPDBM into PMMA film.⁴³ The PL spectra of TPDBM-doped PMMA film in N₂/O₂ mixed gas with different volume ratios were investigated to roughly evaluate the oxygen sensitivity of this PMMA film. As shown in Fig. 8a, totally different PL spectra were obtained at different O₂ concentrations. In comparison to the PL spectra in pure N₂, although the O₂ concentration was as low as 0.1%, a slight but obvious change was still observed at around 500 nm. Also, this emission gradually rose as the O₂ concentration increased. This result was caused by the emergence of photochemical products in the PMMA film, according to the PL spectra of TPBMZ-doped PMMA film. Thus, the photochemical reaction of TPDBM-doped PMMA film could be applied as an efficient and hypotoxic fluorescent probe, with changes in the film fluorescence being used to detect oxygen. The lowest detectable oxygen volume ratio using this film was as low as 0.1%, which was almost the best result compared to most other probes based on phosphorescence quenching theory. In addition, for the same reason, it is also easy to observe the UV-induced fluorescence tuning of TPDBM solid

powder (Fig. 8b). When the blue-emitting TPDBM solid powder was ground under 365 nm UV-irradiation in air, its fluorescence rapidly turned to relatively weaker yellow fluorescence, while under pure N₂ conditions, it remained unchanged. It is worth noting that this phenomenon was also reported by Anu Kundu *et al.*, but they mistakenly thought that no chemical change occurred during this process.⁴¹ This demonstrates that TPDBM has great application potential in oxygen sensors and related photoelectrical areas as a result of the photochemical reaction.

Experimental section

Materials and instruments

All of the reagents and solvents were purchased from commercial sources and used directly without further purification. ¹H NMR spectra of the synthesized compounds were recorded using a Bruker AVANCE III instrument (Bruker, Switzerland). Mass spectra (MALDI-TOF-MS) analysis was performed using an AXIMA-CFRTM-plus instrument. Photophysical properties were investigated using a Shimadzu UV-1800 spectrophotometer (Shimadzu, Japan) and a HORIBA fluorolog-3 luminescence spectrophotometer (HORIBA, Japan). X-ray diffraction studies were carried out using an X'Pert Pro instrument and the single crystal structure was analysed using an Agilent Gemini instrument. Spin coating was carried out with a WS-650MZ-8NPPB spin coater. The UV

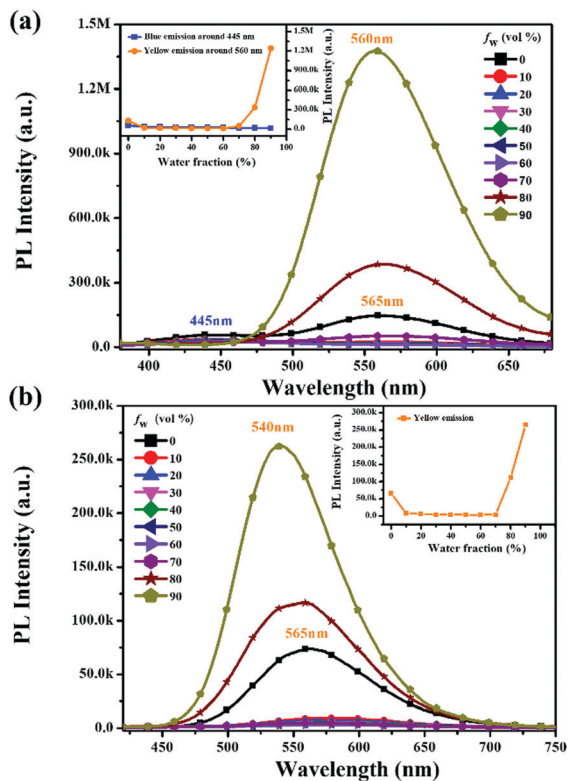


Fig. 7 The PL spectra of (a) TPBMZ and (b) TPDYZ in THF/H₂O mixed solution with different water fraction (f_w , v/v) levels; line charts showing the maximum PL peak intensities of TPBMZ and TPDYZ, respectively, are embedded in each part.

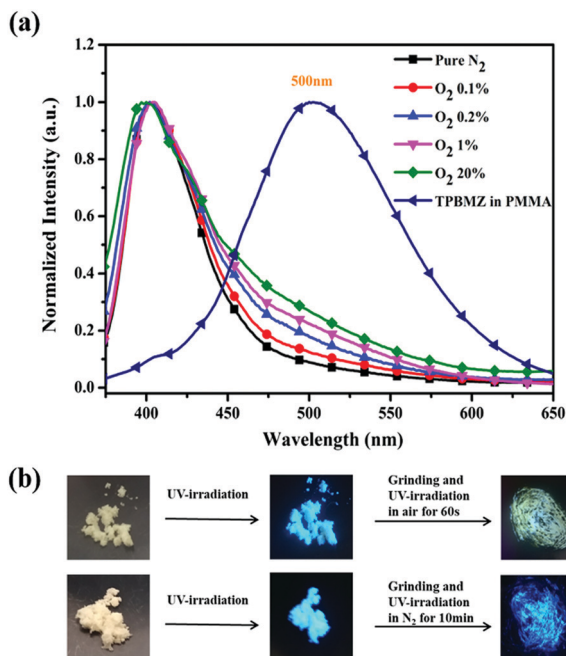


Fig. 8 (a) Normalized PL spectra of TPDBM doped PMMA film in N₂/O₂ mixed gas with different volume ratios. (b) Images of TPDBM solid powder under different treatment conditions.

light used in the photochemical reaction was generated using a common ZF-7 portable UV light (6 W, 365 nm).

For oxygen detection: the ratios of N₂ and O₂ were controlled, respectively, through adjusting gas flows using a gas flowmeter. N₂ and O₂ were thorough mixed until stable in a big container before measurements were taken. The speeds of the gas flows of N₂ and O₂ were controlled at 1000:1, 500:1, and 100:1 ml min⁻¹, respectively, to get ratios of about 0.1%, 0.2%, and 1%. The oxygen probe TPDBM was doped into PMMA film. PL spectrum measurements for oxygen detection were carried out using a HORIBA fluorolog-3 instrument; the excitation wavelength was 350 nm, the emission wavelength range was 360 nm to 680 nm, the slit width was 5 nm, and the integration time was 1 s. The irradiation all came from the excitation light source and, thus, the duration of irradiation for every measurement was about 320 seconds.

Synthesis

Synthesis of 4,4'-(phenylazanediy) dibenzaldehyde (M1). In a round-bottom flask, 9.2 ml of POCl₃ was added to dry DMF (11.2 ml) drop-by-drop at 0 °C. After stirring for 45 minutes, triphenylamine (1.96 g, 8 mmol) was added into the mixture and stirring was continued until the colour became yellow. Then, the reaction mixture was heated and stirred at 65 °C for 16 h. After cooling to room temperature, the reaction mixture was poured into an ice-bath and neutralized with NaOH followed by extraction with ethyl acetate several times. The extracted organic phase was dried using anhydrous MgSO₄ and filtered to obtain the crude product M1. Finally, M1 was further purified *via* column chromatography using petroleum ether/ethyl acetate (10/1) as the eluent to afford a faint yellow solid powder (1.9 g, 80%). ¹H NMR (500 MHz, DMSO, ppm): δ = 9.89 (s, 2H), 7.85 (d, J = 8.5 Hz, 4H), 7.48 (t, J = 7.8 Hz, 2H), 7.32 (t, J = 7.4 Hz, 1H), 7.22 (d, J = 7.8 Hz, 2H), 7.18 (d, J = 8.5 Hz, 4H).

Synthesis of N-phenyl-4-(1,4,5-triphenyl-1H-imidazol-2-yl)-N-(4-(1,4,5-triphenyl-1H-imidazol-2-yl)phenyl)aniline (TPDBM). A mixture of 4,4'-(phenylazanediy) dibenzaldehyde (0.5 g, 1.65 mmol), aniline (1.5 g, 16 mmol), benzil (0.87 g, 4.1 mmol), ammonium acetate (1.27 g, 16 mmol) and acetic acid (30 ml) was refluxed under a nitrogen atmosphere in an oil bath for 24 h. After cooling to room temperature, 200 ml of water was added into the mixture. Then, the solid product was filtered out and washed with deionized water several times. Finally, it was further purified *via* column chromatography using dichloromethane/ethyl acetate (80/3) as the eluent to afford a white solid powder (0.84 g, 60%). ¹H NMR (500 MHz, CD₂Cl₂, ppm): δ = 7.60 (d, J = 6.3 Hz, 4H), 6.90 (d, J = 8.8 Hz, 1H), 7.46–7.23 (m, 24H), 7.22–7.17 (m, 4H), 7.17–7.10 (m, 7H), 6.95 (d, J = 8.5 Hz, 4H). HRMS (mass, m/z): 836.3 [M⁺], C₆₀H₄₃N₅ calc.: 834.04.

Preparation of TPBMZ and TPDYZ. TPDBM (0.5 g, 0.6 mmol) was dissolved in toluene and irradiated *via* a 365 nm UV lamp for 24 h. After rotary evaporation, the resulting solution was purified *via* column chromatography using petroleum ether/ethyl acetate (3/1) as the eluent. Two forms of yellow solid powder, TPBMZ (200 mg, 40%) and TPDYZ (100 mg, 20%), were obtained simultaneously. TPBMZ: ¹H NMR (500 MHz, DMSO, ppm): δ = 7.63 (d, J = 8.8 Hz, 2H), 7.60 (d, J = 7.1 Hz, 2H), 7.57

(d, $J = 7.0$ Hz, 2H), 7.50 (d, $J = 4.4$ Hz, 1H), 7.49 (d, $J = 7.1$ Hz, 2H), 7.43–7.15 (m, 27H), 7.08 (t, $J = 7.2$ Hz, 1H), 7.05 (d, $J = 7.5$ Hz, 2H), 6.88 (d, $J = 8.8$ Hz, 2H), 6.78 (d, $J = 8.9$ Hz, 2H). HRMS (mass, m/z): 867.9 [M^+], $C_{60}H_{43}N_5O_2$ calc.: 866.04.

TPDYZ: 1H NMR (500 MHz, DMSO, ppm): $\delta = 7.67$ (d, $J = 8.8$ Hz, 4H), 7.60 (dd, $J = 9.6, 8.3$ Hz, 8H), 7.51 (t, $J = 7.4$ Hz, 2H), 7.41 (dd, $J = 15.0, 7.4$ Hz, 4H), 7.32 (m, 12H), 7.26–7.20 (m, 5H), 7.09 (t, $J = 7.4$ Hz, 2H), 7.00 (d, $J = 7.3$ Hz, 2H), 6.83 (d, $J = 8.8$ Hz, 4H).

Preparation of PMMA film doped with TPDBM. PMMA powder (0.3 g) was dissolved in acetone (30 ml) followed by the addition of TPDBM or TPBMZ (10^{-4} M) in acetone solution. The mixed solution was heated at 60 °C for 30 minutes and then decanted onto a Petri dish, which was totally dried in a drying oven at 60 °C to obtain the doped film.

Conclusions

In summary, a new triphenylamine-imidazole molecule TPDBM has been designed and successfully synthesized, which exhibited an obviously large red-shifted aggregation-state emission peak at around 520 nm. This unusual red-shifted emission was finally proved to arise from emission from a small number of products produced by a photochemical reaction involving TPDBM through an efficient Förster resonance energy transfer process, rather than from excimer emission. In addition, both of the corresponding photochemical products TPBMZ and TPDYZ exhibited excellent AIE properties, and TPBMZ could be further turned to TPDYZ due to the quite fast reaction rate. Furthermore, the photochemical reaction of TPDBM doped in an operable and hypotoxic PMMA film was employed as a sensitive fluorescent probe for oxygen with a lowest detectable oxygen volume ratio of 0.1%. In this work, the actual origin of the newly apparent unusual red-shifted aggregation-state emission from an imidazole molecule has been confirmed, and this will make a contribution to basic photochemical reaction studies of imidazole derivatives. In addition, the introduction of this molecular configuration can also result in new material systems for sensitive fluorescent probes for oxygen.

Conflicts of interest

There are no conflicts to declare.

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