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Particle size effects in the kinetic trapping of a structurally-locked form of a flexible MOF†

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The application of metal–organic frameworks (MOFs) for gas storage, molecular separations and catalysis necessitates careful consideration of the particle size and structuralisation (e.g. pelletisation, surface-anchoring) of a material. Recently, particle size has been shown to dramatically alter the physical and structural properties of certain MOFs, but overall there is limited information on how the particle size affects the properties of flexible MOFs. Here we demonstrate that the particle size of a flexible MOF, specifically the as-synthesised form of [Cu(bc₂ppm)H₂O]·S (H₂bc₂ppm = bis(4-(4-carboxyphenyl)-1H-pyrazolyl)methane, S = solvent) (**1**), correlates with the rate of structural reorganisation from a “kinetically-trapped”, activated 3D form of this MOF to an “open” 2D form of the structure. We also outline two methods for synthetically reducing the particle size of **1** at room temperature, using 0.1 M NaOH (for two reaction times: 0.5 and 16 h) and with the sodium salt of the ligand Na₂bc₂ppm, producing crystals of 85 ± 15, 280 ± 14 and 402 ± 41 nm, respectively.

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Introduction

Metal–organic frameworks (MOFs) are versatile, highly porous materials whose pore size and functionality can be dictated by the organic ligands (linkers) and metal centres (nodes) that form the framework.^{1–3} MOFs exhibiting structural flexibility represent a subset of these materials, and feature structural changes to their ordered coordination networks (crystal-to-crystal transformations), often reversibly and without changes in the bond sequences within the lattice.^{4,5} This includes transformations induced by host–guest interactions,^{6,7} and physical stimuli such as pressure,⁸ light, and heat.^{9–11} The properties associated with flexible MOFs have led to potentially diverse applications for these materials, including

molecular separations,¹² gas adsorption,^{13–16} sensing,¹⁷ drug delivery (guest capture and release)^{18,19} and catalysis.^{5,20,21}

MOFs are conventionally synthesised under laboratory conditions that consistently produce micron size crystals, primarily due to ease of structural characterisation. While this is adequate for the discovery of new materials, the development of synthetic protocols^{22–26} that also provide control over the crystal size (ideally in the nm range) and morphology is crucial if MOFs are to be tailored for specific applications.²⁷ Such methods are also expected to be crucial if MOFs are to be used in tandem with other materials, such as the solid phase additive within mixed matrix membranes (MMMs).^{28,29}

The particle size of MOFs has also been demonstrated to impart unique physical and structural properties to the materials when precisely controlled in the nanoscale regime.^{30–33} Fundamentally, changes in the particle size will directly alter diffusion rates and surface-to-volume ratios, but more sophisticated changes have been observed.^{30–33} One recent example reported that the crystal downsizing of a two-fold interpenetrated MOF exhibiting elastic flexibility (between an open and closed phase) resulted in the evolution of a third, metastable phase that affords a “shape memory” effect not seen in larger, micrometre-sized crystals.³² In general, however, studies addressing the effect of particle size on the dynamic properties of MOFs (*i.e.* framework flexibility) remain rare, and more studies of this type are required prior to consideration of these materials for use in real-world applications. Herein, we specifically investigate the impact of

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† Electronic supplementary information (ESI) available: Unit cell parameters, combined descriptions of the single crystal structures of [Cu(bc₂ppm)H₂O]·S (S = as-synthesised, acetone, water, ethanol; heated sample), SEM images, dynamic light scattering data, PXRD data, UV-vis transmittance data, gas adsorption data and additional synthetic data. CCDC 1451302 and 1451303. For ESI and crystallographic data in CIF or other electronic format see DOI: 10.1039/c6ce00082g

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particle size on a structural transformation that is controlled by inter-framework chemical bond formation.

We have previously reported a flexible copper(II) MOF, [Cu(**bcppm**)H₂O]·S (1·S, where H₂**bcppm** = bis(4-(4-carboxyphenyl)-1H-pyrazolyl)methane, S = solvent), which has exceptional CO₂/N₂ selectivity when activated.³⁴ This MOF undergoes an intriguing structural transformation from a flexible 2D layered material to a “locked” 3D framework upon activation.³⁴ The role of this structural “locking” is of significant importance to the properties of the framework, and herein we show that this is a reversible particle size-dependent process. We report two methods of synthesising 1·S at room temperature (RT) which afford a dramatic reduction in particle size. The first utilises a strong base to promote rapid deprotonation of the dicarboxylic acid H₂**bcppm**, inspired by the pH control achieved by Zhang *et al.*³⁵ The second uses the sodium salt of the ligand, Na₂**bcppm**, in MOF synthesis, drawing from work recently published by Díaz *et al.*³⁶ We report structural flexibility of the as-synthesised 2D form of the framework and show that particle size correlates with the rate of structural reorganisation of 1 – a ‘kinetically trapped’ form – to the ‘open’ material, 1·S.

Experimental

General considerations

All materials were purchased commercially unless otherwise stated. The ligand bis(4-(4-carboxyphenyl)-1H-pyrazolyl)-methane (H₂**bcppm**) was prepared according to literature methods.³⁴ Micron-scale crystals of [Cu(**bcppm**)H₂O]·S (1·S; S = solvent) and the activated form, [Cu(**bcppm**)H₂O] (1) were also prepared by the procedure of Bloch *et al.*³⁴ The solvated samples reported were obtained by soaking samples of 1·S in millilitre volumes of various solvents (acetone, ethanol or water) overnight, or in the case of the heated material by heating a sample of 1·ethanol on a microscope slide at 85 °C for 1 h. Details of the single crystal X-ray structure determinations of the solvent-exchanged forms of 1 are given in the ESI† (Table S1).

Samples for dynamic light scattering (DLS) experiments using a Malvern Zetasizer Nano series (λ = 671 nm) were prepared by suspending each material in excess methanol. A few drops (less than 5) of each solution were added to a cuvette containing methanol (1 ml), and this solution was mixed and sonicated for 5 minutes prior to the DLS experiment.

Scanning electron microscopy (SEM) images were collected on a Philips XL30 scanning electron microscope at Adelaide Microscopy. Samples were dry-loaded onto adhesive carbon tabs on aluminium stubs and coated with carbon.

Solid-state UV/vis transmittance spectroscopy was recorded on a Varian Cary 5000 UV/vis spectrometer with Harrick Praying Mantis DRP accessory. Dried samples were loaded (<10 mg) and packed flat on the plate. %T spectra were collected between 300 and 800 nm. Infrared spectra

were collected on a Perkin-Elmer Spectrum 100 using a universal attenuated total reflectance (UATR) sampling accessory.

Gas adsorption isotherm measurements were performed on a Micromeritics ASAP 2020 Physisorption Analyser. UHP grade (99.999%) N₂ and CO₂ were used for all measurements. The temperatures were maintained at 77 K (liquid nitrogen bath) and 293 K (room temperature water bath).

Powder X-ray diffraction data were collected on a Bruker D8-Advanced X-ray powder diffractometer (parallel X-ray, capillary-loaded) using a Cu K α (λ = 1.5418 Å) radiation source. Samples were mounted in 0.5 mm glass capillaries and data collected for between 2θ of 2° to 52.94° with Φ rotation at 20 rotations per min at 1 second exposure per step at 5001 steps. The data were then converted into xye format and background-subtracted using WinPlotr 2000 software.³⁷ Simulated powder X-ray diffraction patterns were generated from the single crystal X-ray data using Mercury 3.5.1.^{38,39}

Synthesis of materials

Synthesis of Na₂bcppm**.** H₂**bcppm** (0.5 g, 1.28 mmol) and NaOH (1.3 ml, 2 M) were stirred until the solution became clear. The solvent was removed under reduced pressure to give a white solid Na₂**bcppm** (0.532 g, 96%). M.P.: >260 °C. ¹H NMR (500 MHz/DMSO): δ 6.40 (s, 2H, CH₂), 7.46 (d, 4H, H2'/H6'), 7.80 (d, 4H, H3'/H5'), 7.95 (s, 2H, H3 or H5), 8.39 (s, 2H, H3 or H5). ¹³C NMR (500 MHz/DMSO): δ 66.0, 121.0, 125.0, 128.0, 128.7, 130.0, 136.5, 138.7, 170.0.

General room temperature synthesis of [Cu(bcppm**)H₂O]·S (S = solvent) using the base NaOH (1b·S, where b = base).** In a 20 ml screw cap vial, Cu(NO₃)₂·2.5H₂O (36.0 mg, 0.16 mmol) and H₂**bcppm** (41.5 mg, 0.11 mmol) were dissolved in a mixture of DMF (2 ml) and ethanol (0.5 ml). Sodium hydroxide 0.1 M (0.5 ml) was injected into the stirred solution. The resultant solution was stirred for a period of time (0.5–16 hours) at room temperature. The product was washed with DMF (×2) and methanol (×5) to give a light blue microcrystalline material (8 mg, 16%).

General room temperature synthesis of [Cu(bcppm**)H₂O]·S (S = solvent) using the ligand salt Na₂**bcppm** (1s·S, where s = salt).** In a 20 ml screw cap vial, Cu(NO₃)₂·2.5H₂O (36.0 mg, 0.16 mmol) dissolved in DMF (2 ml) and in a separate vial, Na₂**bcppm** (44.1 mg, 0.11 mmol) was dissolved in a mixture of ethanol (0.5 ml) and water (0.5 ml). The ligand solution was injected into the stirred copper(II) solution and the resultant solution was stirred for a period of time (0.5–16 hours). The product was washed with DMF (×2) and methanol (×5) resulting in a light blue microcrystalline material (23 mg, 46%).

Activation procedure. The as-synthesised samples were washed with DMF (×2) and acetone (×7), dried under a nitrogen flow, and then heated under a high vacuum at 120 °C for 1 hour. Activation yielded light green microcrystalline powders for both samples 1b and 1s.

Results and discussion

Structural flexibility of 1

Single-crystals of $[\text{Cu}(\text{bcppm})\text{H}_2\text{O}]\cdot\text{S}$ (**1-S**) prepared by the conventional solvothermal method³⁴ were initially used for a series of single-crystal X-ray structure determinations. The structural flexibility of the as-synthesised 2D form **1** was probed by a series of single crystal X-ray structure determinations on solvent exchanged forms and also for a heated sample, which affords a structure consistent with the 'locked' 3D phase obtained following activation (Fig. 1). All structure determinations confirmed the formulation of the MOF as reported previously for $[\text{Cu}(\text{bcppm})\text{H}_2\text{O}]\cdot\text{S}$ (ref. 34) but with the changes in the unit cell parameters, described below, that represent the breathing of the MOF. For completeness, structure descriptions are provided in the ESI† but only the unit cell changes are discussed here. In all cases, the samples underwent structural changes in a single-crystal to single-crystal process. The solid-state structure of **1**-acetone revealed a subtle increase in the *a* axis (along the pore channel direction), an increase in the *b* axis, and a decrease in the *c* axis, relative to as-synthesised **1-S** (containing predominantly DMF in the voids, Fig. 1a).³⁴ Similar changes to the *b* and *c* axes were observed for the water- and ethanol-exchanged samples, although in these cases, a reduction of the *a* axis was observed (Fig. 1a and Table S2†). This small *a*-axis contraction for **1**-ethanol and **1**-water corresponds to a reduction of the interlayer separation and is indicative of closer contact between the water ligands bound to the Cu centres of one layer

and the Cu atoms in an adjacent 2D layer, en-route to the formation of the bridging water ligand as identified in the 3D 'locked' form.³⁴ The most dramatic changes in these samples occurred in the *b* and *c* axes, resembling a trellis-type rearrangement (Fig. 1b and c) and an overall structural contraction as indicated by the cell volume. This indicates that the MOF is indeed flexible and responds structurally to guest exchange along all three axes.

As we reported previously,³⁴ an even more dramatic change was observed when **1**-ethanol was heated at 85 °C for 1 hour (slightly lower than used to obtain the activated form), yielding a partially desolvated sample. This form has a significantly reduced cell volume, due to a prominent trellis-like contraction of the *c*-axis and elongation of the *b*-axis. In contrast to the solvent exchanged forms, a considerable contraction of the *a*-axis was observed in this partially desolvated sample corresponding to a change in the coordination environment of the copper(II) centre from five coordinate to a Jahn–Teller distorted octahedral six coordinate. The water ligand bridges the 2-D layers ($\text{Cu1-O1} = 2.265 \text{ \AA}$, $\text{O1-Cu1a} = 2.397 \text{ \AA}$, symmetry code: $1/2 + x, 1/2 - y, 3/2z$) to form the 'locked' 3D structure.

Particle size control of 1

Having established the structural flexibility of **1**, we investigated methods to systematically vary the particle size. Solvothermal synthesis of **1** by the literature procedure,³⁴ produced large, dark-blue block-shaped crystals approximately 50–100 μm in size (Fig. 2d). Since the formation of the framework must necessarily proceed *via* deprotonation of the carboxylate groups, we surmised that increasing the pH of the reaction mixture would promote rapid deprotonation of carboxylate ligands, facilitating faster nucleation and a greater number of nucleation sites to give smaller crystals.^{30,35} Indeed, the use of basic conditions (0.1 M NaOH) over 16 h at room temperature afforded uniform crystals of **1b-S** (where *b*

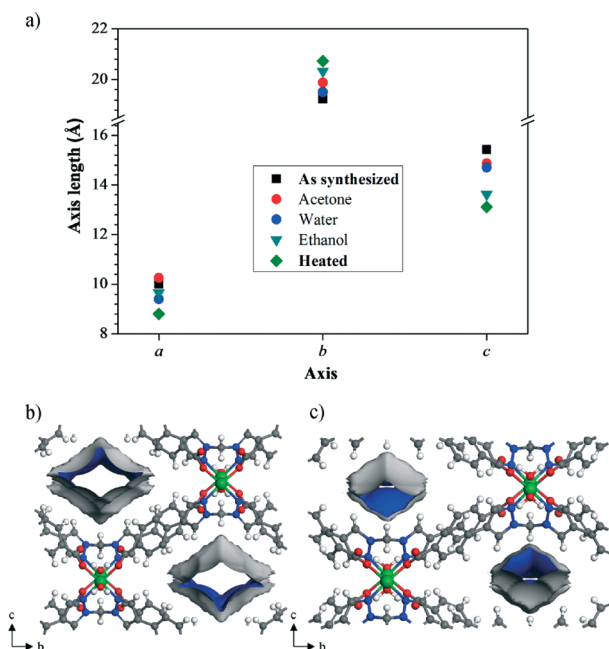


Fig. 1 a) A plot of the crystallographic cell parameters of solvent exchanged samples of **1**. b) A portion of the crystal structure of as-synthesised **1-S** and c) after heating and partial desolvation at 85 °C. Green, gray, blue, red, and white spheres represent Cu, C, N, O, and H atoms, respectively. The contours in the diagram represent the solvent-accessible pore surface as probed with a 3.3 Å sphere.

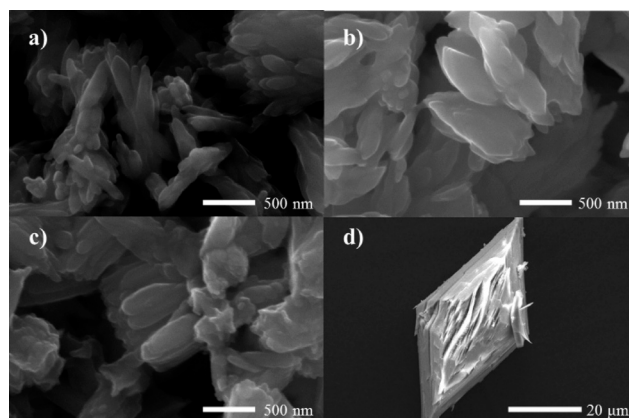


Fig. 2 SEM images of crystals of **1** prepared at room temperature using a reaction solution containing 0.1 M NaOH over a) 0.5 h and b) 16 h, or c) with Na_2bcppm over 16 h, and d) a conventional preparation over 16 h.

= base) approximately 200–300 nm in size (Fig. 2b). Under similar conditions, using the sodium salt of the dicarboxylic acid, Na_2bcppm , produced crystals of **1s-S** (where s = salt) approximately 500 nm in size (Fig. 2c, see also Fig. S3†). Powder X-ray diffraction (Fig. S6†) confirmed that both **1b-S** and **1s-S** were phase-pure, while dynamic light scattering (DLS) experiments revealed that the basic conditions produced a narrower distribution of crystal sizes than the ligand salt conditions, with average sizes of 280 ± 14 nm and 402 ± 41 nm, respectively (Fig. S4†).

Experiments conducted using variable reaction times demonstrated that nucleation occurs rapidly (<0.5 h), as shown by a DLS measurement conducted after 30 minutes (85 ± 15 nm, see Fig. 2a, Fig. S5†). Beyond the initial nucleation events (>0.5 h), there is a period of continuing crystal growth, shown by the broad distributions of particle size at 1.0 and 1.5 h, before particle size plateaus from 1.5 h, likely resulting from limited remaining (deprotonated) ligand in solution. In relation to these observations there are two potential mechanisms of crystal growth at room temperature which have been described in the literature, (a) fast initial nucleation followed by slow crystal growth and (b) slow nucleation followed by rapid crystal growth.^{35,40} Mechanism (a) results in a narrow distribution of particle sizes, whereas (b) results in a broad distribution of particle sizes.^{35,40,41} However, these studies utilised capping agents, such as sodium acetate, to modulate crystal growth after initial deprotonation.²⁵ The time course experiments conducted here did not utilise a capping agent and so the growth rate was not modulated. As such, a broader distribution of crystal size was observed during the growth phase, consistent with rapid uninhibited crystal growth.

The reactant concentration and ligand-to-base ratio were also systematically varied to establish their effect on the particle size of **1** (Fig. 3). Here, three levels were selected for the concentration (12 mM, 18 mM and 36 mM in the H_2bcppm ligand) and the ligand-to-base ratio (2:1, 1:1 and 1:2),

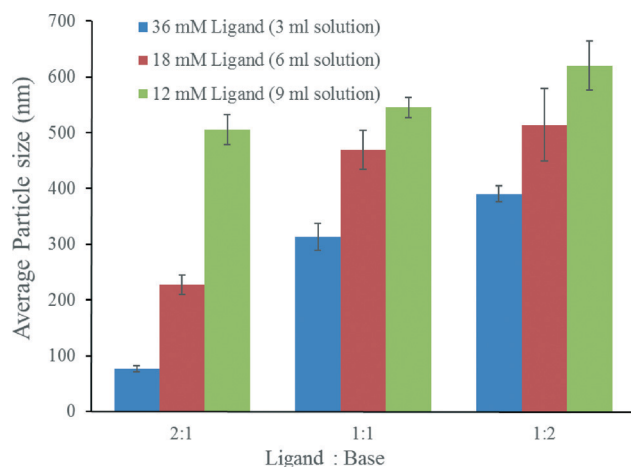


Fig. 3 Average particle size of **1b-S** after 0.5 h by altering the concentration of H_2bcppm and the ligand to base ratio from 2:1 to 1:1 and 1:2.

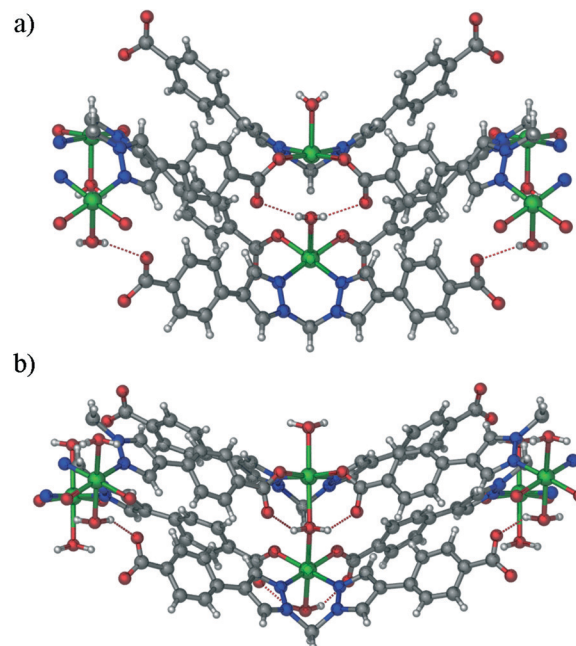


Fig. 4 The coordination environment of copper(II) centres in a) **1-S** and b) **1**, where the dotted red lines denote hydrogen bonding.³⁴ Green, grey, blue, red and white spheres represent Cu, C, N, O, and H atoms, respectively.

giving a total of nine samples for analysis. Isolation of the solids that emerged in each reaction after 0.5 h followed by particle size analysis using DLS revealed two main trends (see Fig. 3). Firstly, the particle size increased when the amount of base was increased (ligand-to-base ratio from 2:1 to 1:2). Secondly, particle size increased when the reactant concentration was reduced. The nine entries in Fig. 3 were each verified to be **1** by PXRD (Fig. S7 and Table S3†). These experiments revealed that reducing the concentration of reactants, from 36 mM of ligand (the limiting reagent) to 12 mM, promoted the growth of larger crystals, as seen in all three ligand-to-base ratios (Fig. 3). This observation is consistent with a decrease in the reactant concentration causing the nucleation of crystals to be more limited, hence allowing the growth of larger crystals. This is in contrast to the coordination modulation approach⁴¹ that gives larger crystals when increasing the concentration, although it must be noted that the concentration was decreased by adding solvent rather than using less reagent (convection and other mass transport phenomena may thus also play a role).

Increasing the amount of base appears to have two effects: expediting the rate of nucleation but also accelerating crystal growth. Raising the pH of a reaction mixture has previously been shown to facilitate deprotonation of carboxylate ligands which in turn leads to an increase in nucleation rate.^{30,35} In the context of this work, the ligand-base ratios; 2:1, 1:1 and 1:2 correspond to a mole ratio of carboxylic acid groups to base in solution of 4:1, 2:1 and 1:1 respectively as H_2bcppm is a dicarboxylic acid. Thus increasing the amount of NaOH added to the solution engenders a greater

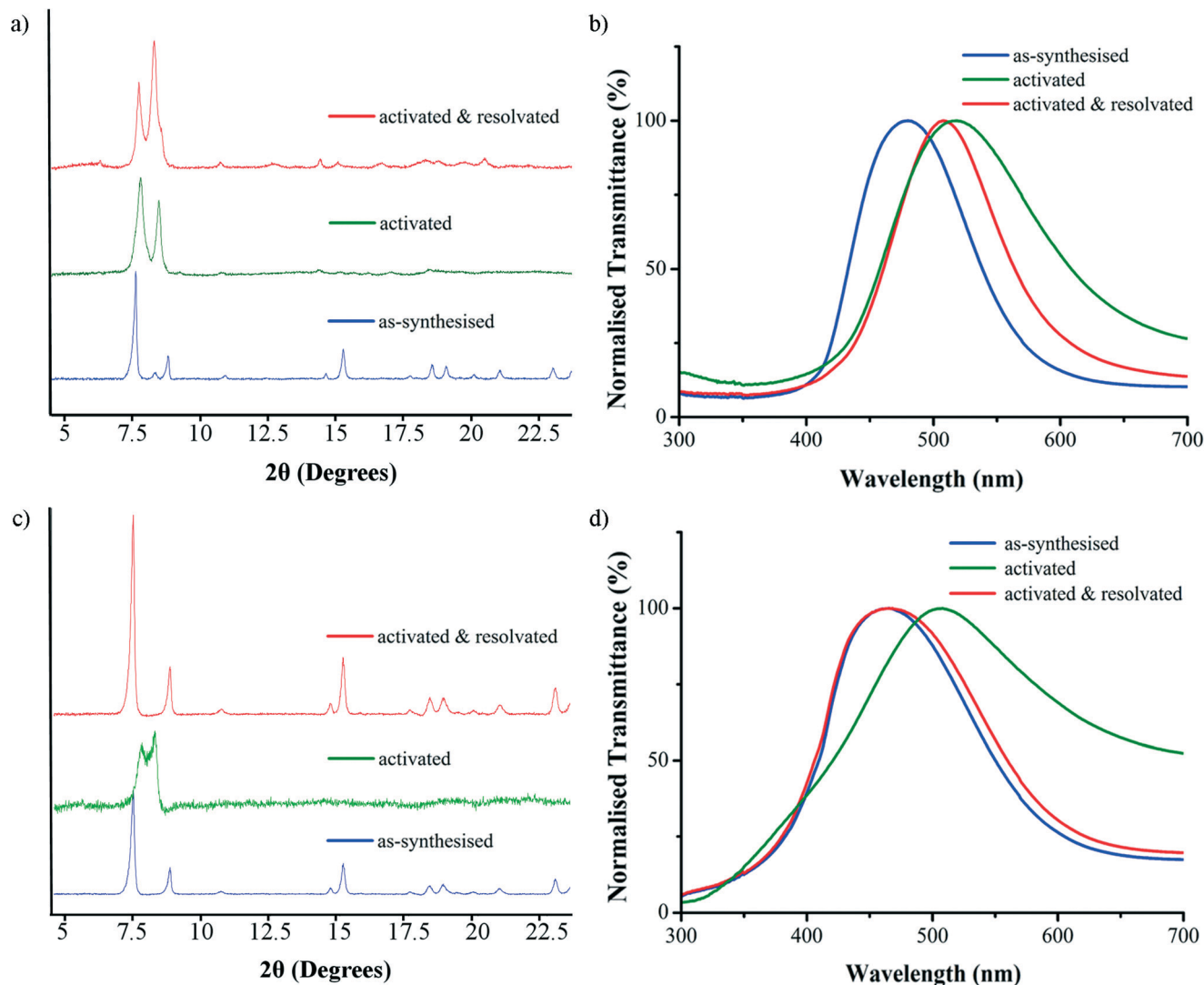


Fig. 5 a) Powder X-ray diffraction patterns and b) solid-state UV/vis transmittance spectra of a methanol-exchanged sample of 1-S (blue), **1** (green) and 1-methanol obtained by resolution of an activated sample of **1** in methanol overnight (red), c) powder X-ray diffraction patterns and d) solid state UV/vis transmittance spectra of a methanol-exchanged sample of 1s-S (blue), **1s** (green) and 1-methanol obtained by resolution of an activated sample of **1s** in methanol overnight (red). In the PXRD plots the peaks at ca. $2\theta = 7.5^\circ$ and $2\theta = 8.5^\circ$ correspond to 011 and 020 planes respectively which reflect the breathing of the MOF. In the sample of 1-S (blue) the peak at ca. $2\theta = 8.0^\circ$ corresponds to a small amount of the dried sample which forms as the material is packed into the capillaries.

concentration of deprotonated ligand in solution and presumably accelerates nucleation. However, while a reduction in crystal size was expected under these conditions, crystal growth has been accelerated as a consequence of having more ligand available in solution and no competition from a suitable modulator. Thus, it appears that the change of parameters predominantly influences the growth process. We suggest that the increase in pH provides more ligand for a limited number of nuclei to grow larger, while the reduction in overall concentration will decrease the number of nuclei formed in the nucleation process and thus gives larger crystals due to the rapid, non-modulated growth process.

Overall, greater control of particle size has been achieved at room temperature with addition of NaOH, affording crystals from 85 nm to 620 nm, by judiciously modifying the

reactant concentration and ligand-to-base ratio. All materials obtained by these varied synthetic methods retain the structural flexibility that is a hallmark of the as-synthesised material.

Kinetic trapping of the 3D locked form of **1**

MOF **1** can be post-synthetically modified by heating a solvent-exchanged sample at 120 °C, thus removing pore bound solvent and activating the material to produce the activated form of **1**. As previously reported, during this process the 2D trellis-like framework 'locks' into a permanently porous 3D framework (Fig. 4),³⁴ and the coordination environment of the copper(II) centre changes from five coordinate square pyramidal to a six coordinate Jahn-Teller distorted

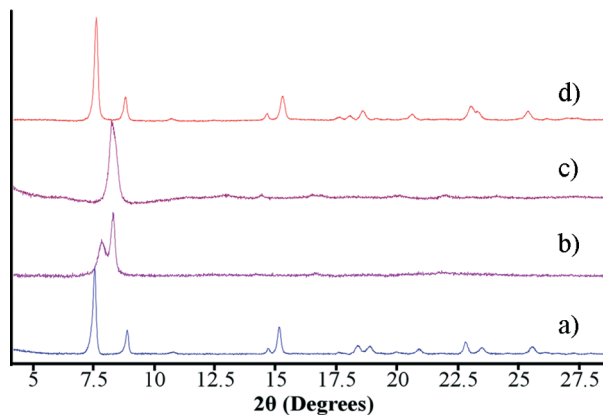


Fig. 6 PXRD of (a) **1b-S** (at 0.5 h), (b) **1b**, and **1b** after immersion in methanol for (c) 1.5 h and (d) 3 h.

octahedral geometry. The apparent ‘locking’ of the framework occurs when the coordinated water ligand on one copper centre coordinates to the vacant site on the square plane of the adjacent centre.³⁴

During activation there is a distinct colour change, from blue to green, that results from a change in coordination environment.³⁴ This distinctive colour change was examined by solid-state UV/vis transmittance spectroscopy, with the maxima shifting from 480 nm (blue) to 500 nm (green), (Fig. 5b). Additionally, the structural change can be observed in the PXRD pattern (Fig. 5a), in which there is a distinct shift in the spacing of the 011 (*ca.* $2\theta = 7.5^\circ$) and 020 (*ca.* $2\theta = 8.5^\circ$) planes. PXRD and UV-vis transmittance spectroscopy both confirm that 50–100 μm sized crystals of **1** could not be ‘unlocked’ by re-solvation in methanol, even when heated at 65 $^\circ\text{C}$.³⁴ By carefully monitoring samples of **1** exposed to methanol it was noted that the 50–100 μm ‘bulk’ crystals took 7 days to undergo a transformation back into the as-synthesised structure (Fig. S8†).

Compared to solvothermally synthesised **1**, decreasing the particle size to 280 ± 14 and 402 ± 41 nm (samples **1b** and **1s**, respectively) increases the rate of structural reorganisation. Upon resolution overnight in methanol, PXRD confirm that both **1s**-methanol and **1b**-methanol (MEOH = methanol) rapidly attain the as-synthesised ‘open’ structure, (Fig. 5c and S9†). In the transmittance spectra there is a notable maxima shift for the resolved material, from 500 nm to 480 nm (Fig. 5d and S9†), indicating a change in coordination environment from six to five coordinate, consistent with the as-synthesised form. The rate of structural reorganisation for sub-micron samples was investigated *via* a series of PXRD experiments. Samples of **1b** (at 0.5 and 16 h) and **1s** (at 16 h), were investigated by collecting PXRD patterns every 90 minutes following immersion in methanol. The submicron samples of **1** demonstrate rapid structural reorganisation and ‘unlocking’ within 3 h of resolution in methanol (Fig. 6, S10 and S11†), which is significantly faster than observed in micron-sized crystals.

Crystal size and the synthesis conditions are the two factors that differ between the bulk and submicron samples. To

eliminate potential influence of the synthetic conditions on the rate of ‘unlocking’ of **1**, a sample of **1** was synthesised using conventional solvothermal conditions (giving large crystals approximately 50–100 μm in size) followed by mechanical grinding to give a sample with a broad range of crystallite sizes from 1 to 10 μm (Fig. S12†). Despite the broad particle size distribution, the as-synthesised structure was re-formed upon resolution in methanol overnight (Fig. S12†). This indicates that the rate of structural unlocking for **1** is not an artefact of the synthetic conditions, but rather originates from a decrease in the particle size.

The diffusion of solvent through the crystal structure is the likely factor mediating structural rearrangement. The kinetic diameter of methanol is 3.6 Å, which is slightly larger than the pore size of **1** (3.58 Å), and diffusion through the pores is likely to be slow at room temperature.^{34,42} Sub-micron sized crystals (280 and 402 nm) of **1** have the same pore size distribution as determined from a fitting to the CO_2 isotherms collected at 273 K (Fig. S13†).^{34,43} Therefore, diffusion of methanol through all samples should occur at the same rate. However, decreasing the average crystal size is expected to reduce the time required for solvation of the crystal and full structural conversion of **1**.

Gas adsorption studies were conducted to assess the porosity of **1**, **1b** and **1s**. **1** is non porous to N_2 at 77 K, but due to increased external surface area, from decreasing crystal size, the sub-micron crystals of **1b** and **1s** yielded isotherms that are best described as type II (Fig. S14†).⁴⁴ In addition, a decrease in overall uptake of CO_2 and a small increase in N_2 uptake at 293 K for **1b** and **1s** afforded a decreased N_2/CO_2 selectivity compared to that previously reported for their micron-sized counterparts³⁴ (Fig. S15 and Table S4†).

Conclusions

The foregoing work has described two synthetic strategies for controlled crystal size reduction of the flexible MOF $\text{Cu}(\text{bcppm})\text{H}_2\text{O}$ (**1**): a room temperature synthesis with 0.1 M NaOH (**1b**, 0.5 and 16 h), and a synthesis using the salt of the linker, Na_2bcppm (**1s**). These strategies both produced sub-micron crystals of 85 ± 15 , 280 ± 14 (**1b** after 0.5 and 16 h), and 402 ± 41 nm for **1s**, respectively. Examining the reaction time revealed a possible mechanism of nucleation under strong basic conditions involving rapid nucleation and unmediated crystal growth from conditions rich in deprotonated ligand. Particle size was further controlled at room temperature with NaOH by varying the reactant concentration and base-to-ligand ratio, affording crystals of different sizes in the range from 85 to 620 nm.

We have shown that the crystal size of **1** correlates with the rate of structural reversibility from the ‘kinetically-trapped form’ to the ‘open form’ of the material. Decreasing the crystal size of **1** reduces the timescale of the ‘unlocking’ step from 7 days for the ‘bulk’ material to within 3 h for sub-micron samples. The increased rate of structural reorganisation presumably originates from more rapid

solvent diffusion into the smaller crystals relative to 'bulk' crystals, due to a larger external surface area of the submicron samples and the shorter distance required for methanol to percolate in the smaller crystals. Thus, particle size has been demonstrated to dramatically alter the physical properties of **1** with regard to structural reorganisation. These observations demonstrate that due consideration must be given before applying MOFs, particularly those with flexible structures, to applications without first examining the properties of the material in the morphology to be utilised.

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References

- (a) H. Furukawa, K. E. Cordova, M. O'Keeffe and O. M. Yaghi, *Science*, 2013, **341**, 974; (b) M. Li, D. Li, M. O'Keeffe and O. M. Yaghi, *Chem. Rev.*, 2014, **114**, 1343–1370.
- (a) O. M. Yaghi, M. O'Keeffe, N. W. Ockwing, H. K. Chae, M. Eddaoudi and J. Kim, *Nature*, 2003, **423**, 705; (b) S. Kitagawa, R. Kitaura and S. Noro, *Angew. Chem., Int. Ed.*, 2004, **43**, 2334; (c) G. Férey, *Chem. Soc. Rev.*, 2008, **37**, 191.
- For recent reviews covering the emerging applications of metal-organic frameworks, refer to the special themed issues in *Chem. Rev.*, and more recently *Chem. Soc. Rev.*: (a) H.-C. Zhou, J. R. Long and O. M. Yaghi, *Chem. Rev.*, 2012, **112**, 673; (b) H.-C. Zhou and S. Kitagawa, *Chem. Soc. Rev.*, 2014, **43**, 5415, and references therein.
- S. Horike, S. Shimomura and S. Kitagawa, *Nat. Chem.*, 2009, **1**, 695–704.
- A. Schneemann, V. Bon, I. Schwedler, I. Senkovska, S. Kaskel and R. A. Fischer, *Chem. Soc. Rev.*, 2014, **43**, 6062–6096.
- P. K. Thallapally, J. Tian, M. R. Kishan, C. A. Fernandez, S. J. Dalgarno, P. B. McGrail, J. E. Warren and J. L. Atwood, *J. Am. Chem. Soc.*, 2008, **130**, 16842–16843.
- S. Henke, A. Schneemann, A. Wutscher and R. A. Fischer, *J. Am. Chem. Soc.*, 2012, **134**, 9464–9474.
- S. A. Moggach, T. D. Bennett and A. K. Cheetham, *Angew. Chem., Int. Ed.*, 2009, **48**, 7087–7089.
- S. Henke, A. Schneemann and R. A. Fischer, *Adv. Funct. Mater.*, 2013, **23**, 5990–5996.
- A. Modrow, D. Zargarani, R. Herges and N. Stock, *Dalton Trans.*, 2011, **40**, 4217–4222.
- Y. Liu, J.-H. Her, A. Dailly, A. J. Ramirez-Cuesta, D. A. Neumann and C. M. Brown, *J. Am. Chem. Soc.*, 2008, **130**, 11813–11818.
- J. R. Li, J. Sculley and H. C. Zhou, *Chem. Rev.*, 2012, **112**, 869–932.
- A. Demessence, D. M. D'Alessandro, M. L. Foo and J. R. Long, *J. Am. Chem. Soc.*, 2009, **131**, 8784–8786.
- J. Liu, P. K. Thallapally, B. P. McGrail, D. R. Brown and J. Liu, *Chem. Soc. Rev.*, 2012, **41**, 2308–2322.
- K. Sumida, D. L. Rogow, J. A. Mason, T. M. McDonald, E. D. Bloch, Z. R. Herm, T. H. Bae and J. R. Long, *Chem. Rev.*, 2012, **112**, 724–781.
- B. Seoane, J. Coronas, I. Gascon, M. Etxeberria Benavides, O. Karvan, J. Caro, F. Kapteijn and J. Gascon, *Chem. Soc. Rev.*, 2015, **44**, 2421–2454.
- Z. Chang, D. H. Yang, J. Xu, T. L. Hu and X. H. Bu, *Adv. Mater.*, 2015, **27**, 5432–5441.
- D. Cunha, M. Ben Yahia, S. Hall, S. R. Miller, H. Chevreau, E. Elkaim, G. Maurin, P. Horcajada and C. Serre, *Chem. Mater.*, 2013, **25**, 2767–2776.
- A. C. McKinlay, J. F. Eubank, S. Wuttke, B. Xiao, P. S. Wheatley, P. Bazin, J. C. Lavalley, M. Daturi, A. Vimont, G. De Weireld, P. Horcajada, C. Serre and R. E. Morris, *Chem. Mater.*, 2013, **25**, 1592–1599.
- K. Uemura, R. Matsuda and S. Kitagawa, *J. Solid State Chem.*, 2005, **178**, 2420–2429.
- S. Li and F. Huo, *Nanoscale*, 2015, **7**, 7482–7501.
- M. Gimeno-Fabra, A. S. Munn, L. A. Stevens, T. C. Drage, D. M. Grant, R. J. Kashtiban, J. Sloan, E. Lester and R. I. Walton, *Chem. Commun.*, 2012, **48**, 10642–10644.
- J. Klinowski, F. A. Paz, P. Silva and J. Rocha, *Dalton Trans.*, 2011, **40**, 321–330.
- A. Pichon, A. Lazuen-Garay and S. L. James, *CrystEngComm*, 2006, **8**, 211.
- T. Friscic, *Chem. Soc. Rev.*, 2012, **41**, 3493–3510.
- S. L. James, C. J. Adams, C. Bolm, D. Braga, P. Collier, T. Friscic, F. Grepioni, K. D. Harris, G. Hyett, W. Jones, A. Krebs, J. Mack, L. Maini, A. G. Orpen, I. P. Parkin, W. C. Shearouse, J. W. Steed and D. C. Waddell, *Chem. Soc. Rev.*, 2012, **41**, 413–447.
- N. Stock and S. Biswas, *Chem. Rev.*, 2012, **112**, 933–969.
- A. Betard and R. A. Fischer, *Chem. Rev.*, 2012, **112**, 1055–1083.
- H. B. Tanh Jeazet, C. Staudt and C. Janiak, *Dalton Trans.*, 2012, **41**, 14003–14027.
- C. Zhang, J. A. Gee, D. S. Sholl and R. P. Lively, *J. Phys. Chem. C*, 2014, **118**, 20727–20733.
- H. Uehara, S. Diring, S. Furukawa, Z. Kalay, M. Tsotsalas, M. Nakahama, K. Hirai, M. Kondo, O. Sakata and S. Kitagawa, *J. Am. Chem. Soc.*, 2011, **133**, 11932–11935.
- Y. Sakata, S. Furukawa, M. Kondo, K. Hirai, N. Horike, Y. Takashima, H. Uehara, N. Louvain, M. Meilikhov, T. Tsuruoka, S. Isoda, W. Kosaka, O. Sakata and S. Kitagawa, *Science*, 2013, **339**, 193–196.
- J. Zhou, X. Yu, X. Fan, X. Wang, H. Li, Y. Zhang, W. Li, J. Zheng, B. Wang and X. Li, *J. Mater. Chem. A*, 2015, **3**, 8272–8275.
- W. M. Bloch, R. Babarao, M. R. Hill, C. J. Doonan and C. J. Sumby, *J. Am. Chem. Soc.*, 2013, **135**, 10441–10448.
- H. Guo, Y. Zhu, S. Wang, S. Su, L. Zhou and H. Zhang, *Chem. Mater.*, 2012, **24**, 444–450.

- 36 M. Sánchez-Sánchez, N. Getachew, K. Díaz, M. Díaz-García, Y. Chebude and I. Díaz, *Green Chem.*, 2015, **17**, 1500–1509.
- 37 T. Rosinel and J. Rodríguez-Carvajal, *Proceedings of the Seventh European Powder Diffraction Conference (EPDIC 7)*, 2000, pp. 118–123.
- 38 C. F. Macrae, P. R. Edgington, E. P. P. McCabe, G. P. Shields, R. Taylor, M. Towler and J. V. D. Streek, *J. Appl. Crystallogr.*, 2006, **39**, 453–457.
- 39 C. F. Macrae, I. J. Bruno, J. A. Chisholm, P. R. Edgington, E. P. P. McCabe, L. Rodríguez-Monge, R. Taylor, J. V. D. Streek and P. A. Wood, *J. Appl. Crystallogr.*, 2008, **41**, 466–470.
- 40 D. Zacher, J. Liu, K. Huber and R. A. Fischer, *Chem. Commun.*, 2009, 1031–1033.
- 41 S. Diring, S. Furukawa, Y. Takashima, T. Tsuruoka and S. Kitagawa, *Chem. Mater.*, 2010, **22**, 4531–4538.
- 42 T. Borjigin, F. Sun, J. Zhang, K. Cai, H. Ren and G. Zhu, *Chem. Commun.*, 2012, **48**, 7613–7615.
- 43 J. Jagiello and M. Thommes, *Carbon*, 2004, **42**, 1227–1232.
- 44 P. B. Balbuenat and K. E. Gubbins, *Langmuir*, 1993, **9**, 1801–1814.