# **Nanoscale**



# **MINIREVIEW**

View Article Online
View Journal | View Issue



Cite this: Nanoscale, 2024, 16, 12329

# Recent advances in synthesis and properties of silver nanoclusters

Xiaolin Liu,†<sup>a,b</sup> Taeyoung Ki,†<sup>a,b</sup> Guocheng Deng,<sup>a,b</sup> Seungwoo Yoo,<sup>a,b</sup> Kangjae Lee,<sup>a,b</sup> Byoung-Hoon Lee,<sup>c</sup> Taeghwan Hyeon (1) \*\*a,b\* and Megalamane S. Bootharaju\*\*<sup>a,b</sup>

Achieving atomic precision in nanostructured materials is essential for comprehending formation mechanisms and elucidating structure—property relationships. Within the realm of nanoscience and technology, atomically precise ligand-protected noble metal nanoclusters (NCs) have emerged as a rapidly expanding area of interest. These clusters manifest quantum confinement-induced optoelectronic, photophysical, and chemical properties, along with remarkable catalytic capabilities. Among coinage metals, silver distinguishes itself for the fabrication of stable nanoclusters, primarily due to its cost-effectiveness compared to gold. This minireview provides an overview of recent advancements since 2020 in synthetic methodologies and ligand selections toward attaining NCs boasting a minimum of two free valence electrons. Additionally, it explores strategies for fine-tuning optical properties. The discussion extends to surface reactivity, elucidating how exposure to ligands, heat, and light induces transformations in size and structure. Of paramount significance are the applications of silver NCs in catalytic reactions for energy and chemical conversion, supplemented by in-depth mechanistic insights. Furthermore, the review delineates challenges and outlines future directions in the NC field, with an eye toward the design of new functional materials and prospective applications in diverse technologies, including optoelectronics, energy conversion, and fine chemical synthesis.

Received 25th April 2024, Accepted 30th May 2024 DOI: 10.1039/d4nr01788a

rsc.li/nanoscale

## 1. Introduction

The use of silver traces back to ancient civilizations like Rome and China, valued for its aesthetic allure and believed medicinal properties. Craftsmen combined silver with other metals for decorative brilliance and employed silver vessels to safeguard liquids from contamination. Over the decades, silver has undergone significant advancements, particularly following Faraday's preparation of colloidal gold and its catalytic application by Haruta, which redirected focus towards silver. Turkevich's synthesis of analogous silver colloids and Brust's pioneering use of thiolate ligands for surface functionalization and size control of gold nanoparticles further contributed to this evolution. Subsequently, researchers developed synthetic methods for producing atomically monodispersed gold and silver nanoparticles, commonly referred to as atomically

precise nanoparticles or nanoclusters (NCs), typically with diameters below 3 nm.<sup>4,5</sup> These NCs, bridging the gap between plasmonic nanoparticles and molecular precursors, offer insights into the evolution of material properties at discrete atom levels.<sup>6-10</sup> Due to quantum confinement effects, NCs exhibit molecule-like optical, photophysical, and chiroptical properties.<sup>11-18</sup> Their high surface-to-volume ratio and defined metal–ligand interface make NCs ideal model catalysts for unravelling molecular-level insights into the origins of catalysis through combined experimental and computational approaches.<sup>19-24</sup>

Among coinage metals, gold has been extensively utilized to investigate the structure–property relationship of gold NCs due to their enhanced stability, facilitating structural characterization *via* single-crystal X-ray crystallography.<sup>1,2</sup> Despite challenges in fabricating stable silver NCs, their lower cost renders them economically competitive, prompting the development of new synthesis strategies.<sup>3</sup> It is noteworthy that silver forms stable Ag(1) clusters with few or no free electrons. Serving as models for metallic nanoparticles, silver NCs with free valence electrons are invaluable for understanding various properties, including identifying the size regime of non-metallic to metallic transitions and emerging plasmonic nanocatalysis.<sup>25–27</sup> Hence, this minireview focuses on the

<sup>&</sup>lt;sup>a</sup>Center for Nanoparticle Research, Institute for Basic Science (IBS), Seoul 08826, Republic of Korea. E-mail: thyeon@snu.ac.kr, msbootharaju@snu.ac.kr

bSchool of Chemical and Biological Engineering, and Institute of Chemical Processes, Seoul National University, Seoul 08826, Republic of Korea

<sup>&</sup>lt;sup>c</sup>KU-KIST Graduate School of Converging Science and Technology, Korea University, Seoul, Republic of Korea

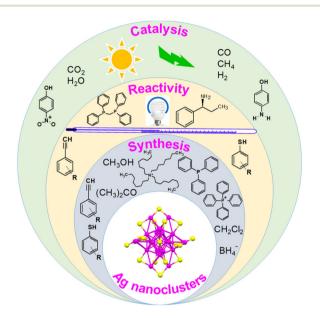
<sup>†</sup>These authors contributed equally.

recent advances in ligand-protected Ag NCs possessing free valence electrons since 2020, covering synthetic methodologies, modulation of optical properties, surface reactivity, and promising catalytic applications (Scheme 1). Additionally, we discuss the challenges and future directions in the field of metal NCs.

#### 2. Synthetic strategies

#### 2.1. Ligand-controlled synthesis

The synthesis of ligand-protected metal NCs is typically carried out through the chemical reduction of metal-ligand precur-



Scheme 1 Schematic illustration of the overall content of the review.

sors in appropriate solvents at low or room temperature. Sodium borohydride is commonly employed as a reducing agent.<sup>28-30</sup> Additionally, arylsilane, cyanoborohydride, and borane alkylamine complex are used to control the rate of reduction of metal-ligand complexes and thereby the size and structure of the resultant NCs. 31-35 Understanding the molecular-level mechanism of the formation of NCs is challenging since the reaction solutions involve different phases, disabling the analysis of reaction intermediates through mass spectrometry. Nevertheless, ligands play a critical role in directing the formation of suitable intermediates, which lead to the formation of NCs of a specific size.

Zhu and co-workers<sup>36</sup> developed a promising synthetic strategy called "dual-level kinetic control", in which the added secondary ligand (phosphine), in addition to the primary thiolate ligand, and the low reaction temperature are found to help achieve first- and second-level kinetic control, respectively. In the absence of a secondary ligand, polydisperse thiolated Ag nanoparticles were obtained. By employing phosphine and diphosphine ligands of different molecular structures, the "dual-level kinetic control" method offers a family of Ag NCs,  $[Ag_{25}(SR)_{18}]^-,$  $[Ag_{34}(SR)_{18}(DPPP)_3Cl_4]^{2+}$  $[Ag_{36}(SR)_{26}S_4]^{2+}$ ,  $[Ag_{37}(SR)_{25}Cl_1]^+$ , and  $[Ag_{52}(SR)_{28}Cl_4]^{2+}$  (Fig. 1) (DPPP: 1,3-bis(diphenylphosphino)propane; SR: 2-adamantanethiolate). In the first-level kinetic control (Fig. 1A), the phosphine ligands retarded the reduction rate, resulting in different Ag NCs correlating with the phosphine ligand type. In the second-level kinetic control (Fig. 1B), lowering the reaction temperature using an ice bath further slowed down the reduction rate, which led to advanced kinetic control and improved synthetic yields of Ag NCs.

By employing various combinations of thiol and diphosphine, a rare collection of super-atomic Ag NCs with sizes Ag<sub>19</sub>, Ag<sub>22</sub>, Ag<sub>26</sub>, and Ag<sub>30</sub>, possessing 2, 4, 6, and 8 free valence electrons respectively, has been reported.<sup>37</sup>

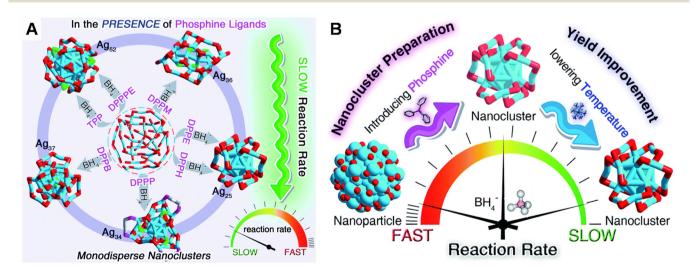


Fig. 1 Schematic illustration of (A) first-level and (B) dual-level kinetic control via introducing phosphine ligands and low reaction temperature, leading to synthesis of a series of Ag NCs with improved yields. Reproduced with permission from ref. 36. Copyright (2022).

Interestingly, the cores of these clusters consist of one decahedral Ag<sub>7</sub>, perpendicular bidecahedra, three-dimensional pentadecahedra, and hexadecahedra, respectively. The strong structural correlation demonstrates the motif-to-core evolution of the surface Ag (on AgS2) to construct additional decahedral blocks, offering opportunities to investigate the impact of decahedral evolution patterns on the properties of NCs. In the presence and absence of phosphine ligands, a dithiol is shown to form Ag NCs of different sizes and structures. Specifically, 1,2-benzenedithiolate (1,2-BDT) forms Ag<sub>14</sub>(1,2-BDT)<sub>6</sub>(PPh<sub>3</sub>)<sub>8</sub> in the presence of  $PPh_3$ , 38 while  $[Ag_9(1,2\text{-BDT})_6]^{3-}$  forms in the absence of the PPh3 ligand. 39 The Ag14 cluster possesses an inner core of Ag<sub>6</sub> arranged in an octahedral shape, surrounded by an outer cubic shell of Ag<sub>8</sub>. Conversely, the Ag atoms of the Ag<sub>9</sub> cluster are arranged in an elongated body-centered cubic format. These examples demonstrate the crucial role of phosphine in the size-selected synthesis of NCs. In contrast, when an isomeric dithiolate, 1,3-benzenedithiolate (1,3-BDT), is used, Ag<sub>29</sub> NCs are synthesized both in the presence and absence of the PPh3 ligand. 40 The formation of the Ag29 cluster is attributed to its geometrical and electronic stability. However, the presence of PPh3 facilitates single-crystal growth and increases the yield of Ag<sub>29</sub> clusters. 40 Notably, the replacement of PPh<sub>3</sub> with diphosphines also yields Ag<sub>29</sub> clusters.<sup>41</sup>

The Zang group synthesized novel Ag(0)-containing Ag NCs by controlling the molecular structures of auxiliary phosphine ligands while keeping the thiolate ligand unchanged. Decifically, they synthesized  $[Ag_{17}(L)_{12}(PPh_3)_6]^{3+}$  (L = quinoline-2-thiol (deprotonated)-4-carboxyl) and  $[Ag_{32}(L)_{18}(P(Ph_2CH_3)_3)_6Cl_2]^{4+}$  (where  $P(Ph-pCH_3)_3 = tri-p$ -tolylphosphine). The  $P(R_3)_3 = tri-p$ -tolylphosphine and  $P(R_3)_3 = tri-p$ -tol

#### 2.2. Counterion- and solvent-controlled synthesis

Counterions (cation or anion) are provided to balance the core charge of the NCs, resulting in electrostatic stabilization. Additionally, the bulkiness of the counterion can direct the formation of NCs with specific size and structure. Zang and co-workers<sup>43</sup> observed the unusual cocrystallization of two anionic Ag NCs,  $[Ag_{62}(MNT)_{24}(TPP)_6]^{8-}$  and  $[Ag_{22}(MNT)_{12}(TPP)_4]^{4-}$  (Fig. 2) (MNT<sup>2-</sup>: dimercaptomaleonitrile; TPP: triphenylphosphine). The inter- and intra-cluster noncovalent interactions drive the cocrystallization of NCs. Interestingly, when the synthesis is carried out in the presence of positively charged counterions with varied bulkiness, a specific sized NC is obtained. Tetrabutylammonium and tetraphenylphosphonium direct the formation of the truncated-tetrahedral  $Ag_{22}$  and the octahedral  $Ag_{62}$  NCs, respectively.

The Kang group obtained three medium-sized Ag NCs, namely  $Ag_{52}$ ,  $Ag_{54}$ , and  $Ag_{58}$ , co-protected by thiolate, halide, and hydride ligands.<sup>44</sup> The  $Ag_{52}$  and  $Ag_{54}$  NCs are synthesized

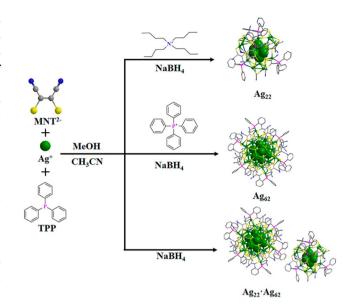


Fig. 2 Synthetic route for counter cation-directed synthesis of  $[Ag_{62}(MNT)_{24}(TPP)_6]^{8-}$  and  $[Ag_{22}(MNT)_{12}(TPP)_4]^{4-}$  NCs. Color legend: green, Ag; pink, P; yellow, S; blue, N; gray, C. Reproduced with permission from ref. 43. Copyright (2023).

in the presence of tetraphenylphosphonium, whereas the  $Ag_{58}$  NC is produced in the presence of tetraoctylammonium counterion. The  $Ag_{52}$  and  $Ag_{54}$  NCs are structurally homologous with respect to the kernel structures and the overall metalatom configurations, while the  $Ag_{58}$  NC has a completely different geometric structure. The formation of different cluster compounds is attributed to the following: (i) the flexibility of the hydrocarbon chain of tetraoctylammonium and the rigidity of the phenyl rings of tetraphenylphosphonium. The different degrees of steric hindrance of the counterions would have resulted in structural rearrangement at distinct levels during the formation of NCs; (ii) different and/or specific interactions among cluster intermediate species and counterions would have remarkably affected the synthesis of NCs of a specific size.

The solvent medium, in which metal-ligand precursors form nuclei and further grow to yield NCs, also plays a crucial role in controlling the size and structure of metal NCs. Organic solvents with varying polarities, such as alcohols (methanol and ethanol), halocarbons (dichloromethane and chloroform), acetonitrile, and dimethylformamide (DMF), are commonly used for the synthesis of Ag NCs since both the precursors and the NC products dissolve well in these solvents. Although the exact reasons for the size selection of NCs are unknown, changing solvents can lead to NCs with distinct structures.

The  $[Ag_{30}(TC4A)_4(TBPMT)_8]$  NCs are synthesized in a methanol-chloroform mixture (where  $TC4A4^- = p$ -tert-butylthiacalix [4]arene; TBPMT = 4-tert-butylbenzenemethanethiolate). Interestingly, by replacing methanol in the solvent mixture with acetone, different-sized NCs  $[Ag_{34}Na_4Cl_4(TC4A)_4(TBPMT)_{11}(solvents)]^{3+}$  are obtained. The 6-electron- $Ag_{14}^{8+}$ 

kernel of Ag<sub>30</sub> is formed by the central Ag<sub>6</sub> octahedron sandwiched between two orthogonally oriented Ag<sub>5</sub> trigonal bipyramids through sharing vertices. On the other hand, the 4-electron-Ag<sub>10</sub><sup>6+</sup> core of Ag<sub>34</sub> is formed by two edge-shared octahedrons. These two clusters exhibit distinct optical absorption profiles due to the differences in their geometric and electronic structures. Similarly, n-butanol, a rarely used alcohol for the synthesis of NCs, is found to be critical in the synthesis of a large Ag NC  $[Ag_{155}(CyS)_{40}(TC4A)_5Cl_2]$  (where CyS = cyclohexanethiolate). 46 The use of other common alcohols is unable to produce this NC. This cluster consists of a metallic core of four concentric shells, Ag<sub>13</sub>@Ag<sub>42</sub>@Ag<sub>30</sub>@Ag<sub>70</sub>. In some instances, special precursors are required to obtain particular clusters. For example,  $[Ag(pz)]_n$  (where Hpz = pyrazole) is necessary to obtain the [Ag<sub>78</sub>(<sup>i</sup>PrPhS)<sub>30</sub>(dppm)<sub>10</sub>Cl<sub>10</sub>]<sup>4+</sup> cluster.47 Here, the pz-ligand is not incorporated into the cluster, but its strong coordination effect serves to make [Ag (pz)<sub>n</sub> a controlled-releasing silver ion source. Such ligandassisted synthesis is applicable to other NCs as well. 48,49 The presence of the halide ligand originates from the in situ decomposition of the halocarbon solvent or other sources, a phenomenon also observed in other cluster systems. 35,47,50,51

#### 2.3. Photochemical and metal-mediated approach

Light serves as a promising energy source for the reduction of noble metal precursors to their metallic state. Inspired by the photochemical synthesis of plasmonic nanoparticles, 52 Zhang and co-workers attempted to synthesize ligand-protected Ag NCs.<sup>53</sup> By irradiating a solution of suitable metal and ligand precursors along with an organic base, they successfully synthesized atomically monodisperse [Ag<sub>25</sub>(4-MePhC=C)<sub>20</sub>  $(Dpppe)_3^{3+}$  NCs (as depicted in Fig. 3A) (where Dpppe = 1,5bis(diphenylphosphino)pentane; 4-MePhC=CH = 4-methylphenylacetylene). The photochemical formation mechanism of the NCs involves a photoinduced electron-transfer (PET) process. Upon photoexcitation of amines with light wavelength <455 nm, electrons from the amine are transferred to Ag<sup>+</sup> ions, resulting in the oxidation of the amine and the formation of amine N-oxide.

The PET process is supported by experimental and density functional theory results. The absence of light and amine leads to no formation of NCs. A minimum photon energy of 2.7 eV is required to initiate the synthesis of NCs. The structure of Ag<sub>25</sub> consists of an Ag<sub>13</sub> core in the form of an anticu-

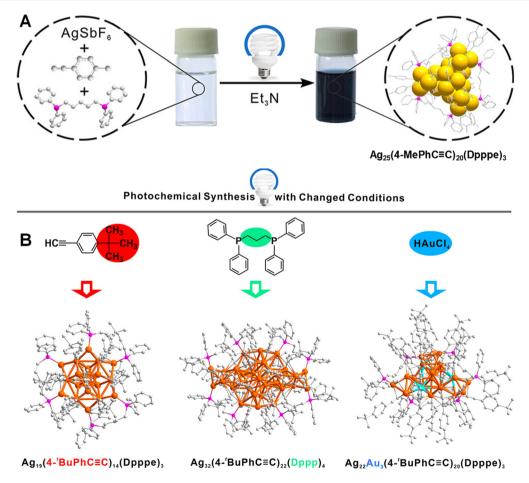


Fig. 3 (A) Schematic illustration of the photochemical synthesis of Ag<sub>25</sub> NCs by irradiating the precursor solution by <455 nm light for 24 h. (B) Schematic of the photochemical synthesis of Ag NCs of different size, structure and composition. Reproduced with permission from ref. 53. Copyright (2023).

boctahedron, a type of hexagonally close-packed (hcp) lattice. This core differs from the Ag<sub>13</sub> icosahedral core of thiolated Ag<sub>25</sub> clusters.<sup>54</sup> This difference likely originates from variations in the coordination modes of thiolate and alkynyl ligands with silver. The remaining 12 Ag atoms form three distorted tetrahedra, which are trigonally distributed around the Ag<sub>13</sub> core. Subsequently, the metal atoms of the cluster are co-protected by diphosphine and alkynyl ligands. The photochemical method has been found to be applicable to the synthesis of Ag NCs of various sizes and structures by using alkynyl and phosphine ligands with different molecular structures and functionalities (as shown in Fig. 3B). Interestingly, by substituting 4-tert-butylphenylacetylene (4-tBuPhC≡CH) for 4-MePhC≡CH,  $[Ag_{19}(4-tBuPhC \equiv C)_{14}(Dpppe)_3]^{3+}$  is synthesized with an overall structure similar to that of Ag<sub>25</sub>. However, the Ag<sub>19</sub> cluster possesses an Ag<sub>13</sub> cuboctahedron core unlike the anticuboctahedron Ag<sub>13</sub> core of the Ag<sub>25</sub> cluster. Additionally, by introducing 1,3-bis(diphenylphosphino)propane (Dppp), the synthesized product contains  $[Ag_{32}(4-tBuPhC = C)_{22}(Dppp)_4]^{3+}$ . The structure of Ag<sub>32</sub> comprises an Ag<sub>24</sub> kernel and two distorted Ag<sub>4</sub> tetrahedra. Interestingly, it exhibits electron paramagnetic resonance (EPR) activity due to its 7 free valence electrons. Furthermore, this photochemical route can be used to synthesize alloy NCs, such as  $[Ag_{22}Au_3(4-tBuPhC \equiv C)_{20}(Dpppe)_3]^{3+}$ , by introducing corresponding metal precursors.

Attempts to synthesize heteroatom-doped Ag NCs sometimes lead to the formation of undoped Ag NCs, which are only possible in the presence of that particular heterometal atom. For instance, Pd is found to mediate the synthesis of Ag<sub>33</sub>(SCH<sub>2</sub>CH<sub>2</sub>Ph)<sub>24</sub>(PPh<sub>3</sub>)<sub>4</sub> NC. <sup>55</sup> In its absence or in the presence of Au or Pt, a different sized NC, Ag<sub>23</sub>(PPh<sub>3</sub>)<sub>8</sub>(SC<sub>2</sub>H<sub>4</sub>Ph)<sub>18</sub>, is obtained. Control experiments indicated that Pd(PPh<sub>3</sub>)<sub>4</sub> is formed first by borohydride reduction, which further catalyzes the reduction of Ag<sup>+</sup> and the growth of the Ag<sub>33</sub> cluster. Similarly, Cu is found to be essential in synthesizing Ag<sub>44</sub>(EBT)<sub>26</sub>(TPP)<sub>4</sub> NC (EBT: 2-ethylbenzenethiolate),<sup>56</sup> whose size and core structure are analogous to an all-thiolated Ag<sub>44</sub>(SR)<sub>30</sub> cluster. The formation of AgCu alloy NCs is believed to mediate the synthesis of electronically and geometrically stable Ag44 clusters. On the other hand, by replacing Cu with Au, AgAu alloy clusters are formed, indicating the unique role of specific metals in the synthesis of certain Ag NCs.

#### 2.4. Other ligands for synthesis of NCs

Coinage metal NCs are typically synthesized using soft-base ligands such as thiolates, phosphines, selenolates, and alkynylides. The use of hard-base ligands with O and N donors is rare because the combination with soft bases is not stable. Recently, the [Ag<sub>8</sub>(pfga)<sub>6</sub>]<sup>6-</sup> cluster (pfga: perfluoroglutarate) was synthesized with an unprecedented rhombohedral Ag<sub>8</sub><sup>6+</sup> core.<sup>57</sup> This two-electron cluster exhibits visible light absorption and emits bright green-yellow light. Following this work, 9-anthracene carboxylate, a bulky monocarboxylate, and different phosphines were used to synthesize two-electron  $[Ag_{16}(L)_8(9\text{-AnCO}_2)_{12}]^{2^+} \ clusters.^{58} \ Its \ structure \ features \ an$  $[Ag_8@Ag_8]^{14+}$  core-shell structure, with an  $[Ag_8]^{6+}$  superatomic inner core adopting an unprecedented structure showing a distorted hexagonal bipyramidal structure. The outer Ag<sub>8</sub> shell is composed of the unique "Ag(PR3)-AnCOO-Ag(PR3)" staple motifs.

Using N donor ligands, Wang and co-workers synthesized  $[Ag_{61}(dpa)_{27}]^{4+}$  (Hdpa = dipyridylamine), the longest silver supercluster based on icosahedral Ag13 building blocks with vertex sharing.<sup>59</sup> In this large cluster of clusters, the icosahedron units are bridged by dpa ligands. The quantum size effects contribute significantly to the enhanced metal-related absorptions and the red-shift in the near-infrared region as the length of the cluster increases from 8-electron [Ag<sub>21</sub>(dpa)<sub>12</sub>]<sup>+</sup> NC to 30-electron  $[Ag_{61}(dpa)_{27}]^{4+}$  NC (Fig. 4). Furthermore, electron coupling is developed between constituent Ag<sub>13</sub> building units. Notably, a disc-like largest structurally characterized Ag cluster of clusters,  $^{31}$   $[Ag_{93}(PPh_3)_6(C \equiv CR)_{50}]^{3+}$  (R =4-CH<sub>3</sub>OC<sub>6</sub>H<sub>4</sub>) is observed for a phosphine and alkynyl combination.

Continuing the use of N donor ligands for Ag NC synthesis, Wang and co-workers<sup>60</sup> designed tetradentate formamidinates and successfully obtained two fcc clusters: [Ag<sub>52</sub>(5-Fdpf)<sub>16</sub>Cl<sub>4</sub>]<sup>2+</sup> (5-F-Hdpf = N,N'-di(5-fluoro-2-pyridinyl)formamidine) and  $[Ag_{53}(5-Me-dpf)_{18}]^{5+}$  (5-Me-Hdpf = N,N'-di(5-methyl-2-pyridinyl)formamidine). All silver atoms are engaged in fcc packing with [111] facets, and no staple motifs are present due to the linear arrangement of the four N donors of the dpf ligands. The Ag<sub>53</sub> cluster is the largest well-defined [111] facet with 14 fcc Ag atoms. In combination with organic ligands, inorganic ligands such as oxometalates are used to stabilize and tune the cluster size. 61,62 For example, 20-electron Ag NCs such as  $Ag_{28}(dppb)_6(MoO_4)_4$ ,  $Ag_{28}(dppb)_6(WO_4)_4$  and  $Ag_{32}(dppb)_{12}$  $(MoO_4)_4(NO_3)_4$  (dppb: 1,4-bis(diphenylphosphino)butane) are synthesized through co-protection. Combination of novel multiple ligands can lead to new assembled clusters. For instance, alkynyl, trifluoroacetate, and diphenylphosphinate set forms luminescent one-dimensional high-nuclearity silver polyclusters, 63 such as  $[Ag_{16}(C \equiv CtBu)_6(tfa)_6(Ph_2PO_2)_4(CH_3OH)]_n$ , and  $[Ag_{21}(C \equiv CtBu)_8(tfa)_6(Ph_2PO_2)_7]_n$ .

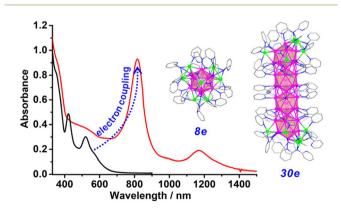


Fig. 4 Optical absorption spectra of [Ag<sub>21</sub>(dpa)<sub>12</sub>]<sup>+</sup> and [Ag<sub>61</sub>(dpa)<sub>27</sub>]<sup>4+</sup> NCs and corresponding atomic-level structures. Reproduced with permission from ref. 59. Copyright (2021).

#### 2.5. Stability, yield, and functionalization

Metal NCs are typically metastable, which limits their potential applications in optoelectronics and catalysis. For example, the weak Ag-P bonding in the 58-electron Ag<sub>64</sub>(P<sup>n</sup>Bu<sub>3</sub>)<sub>16</sub>Cl<sub>6</sub> cluster makes the NC unstable.64 However, the stability of NCs has been significantly improved by employing a combination of ligands. Zhu and co-workers<sup>65</sup> designed a strategy termed "surface environment complication", in which the use of a bidentate PPh<sub>2</sub>py ligand with dual (P and N) coordination sites helps stabilize the extremely unstable Ag<sub>29</sub>(S-Adm)<sub>18</sub>(PPh<sub>3</sub>)<sub>4</sub> cluster by forming a new Ag<sub>29</sub>(S-Adm)<sub>15</sub>(NO<sub>3</sub>)<sub>3</sub>(PPh<sub>2</sub>py)<sub>4</sub> cluster. The Cl intercalation in plasmonic Ag<sub>307</sub>Cl<sub>62</sub>(SPh<sup>t</sup>Bu)<sub>110</sub> cluster leads to its superstability (stable for a week at 55 °C).66 Pradeep and co-workers<sup>67</sup> used carboranethiol for the preparation of propeller-shaped [Ag<sub>21</sub>(MCT)<sub>12</sub>(PPh<sub>3</sub>)<sub>2</sub>]<sup>+</sup> (MCT: m-carborane-9-thiolate). This cluster is found to be remarkably thermally stable up to 100 °C. The use of macrocyclic ligands in combination with thiolate made the 155-atom Ag NC ultrastable.46 It exhibits an outstanding photothermal effect. This cluster can attain a temperature of 59.1 °C within 300 s. The heating ability of the cluster remains nearly intact through 6 cycles of heating and cooling processes.

High yields of NCs are crucial for exploring their properties. Since controlling the monodispersity of clusters in large-scale synthesis is challenging, more efforts are indeed required to achieve satisfactory yields. Fast and high-yield (83%) synthesis of Ag<sub>44</sub>(SR)<sub>30</sub> clusters is achieved through the development of the CTAB (cetyltrimethylammonium bromide) reverse micelle method.<sup>68</sup> Forming reverse micelles using CTAB to construct a sealed chemical environment is found to be critical in achieving fast and high-yield synthesis. Functionalization of the cluster surface can be useful in improving stability and other properties. Tang and co-workers<sup>69</sup> used a thiophenol ligand with silane substituents, which not only enabled the synthesis of  $[Ag_{44}(SPhSi(OEt)_3)_{30}]^{4-}$  clusters but also improved their stability up to 60 °C in air for 13 hours. This silica-coated NCs may be useful in photothermal and photoacoustic imaging applications.

#### 3. Tuning optical properties

The PL of Ag NCs is an attractive characteristic with potential applications in sensing and imaging. It is influenced by several factors, including the metal core, ligand, dopant, solvent, intercluster distance, ion-pairing, and supramolecular interactions, which have been covered in the previous reviews. 70-75 The hydrophilic Ag NCs typically exhibit weak PL due to quenching. To investigate the underlying mechanism of PL quenching in hydrophobic Ag NCs in an aqueous phase, Zhu and colleagues developed a surface functionalization strategy (Fig. 5).<sup>76</sup> Specifically, they synthesized a functionalized [Ag<sub>29</sub>(BDTA)<sub>12</sub>(DPBA)<sub>4</sub>]<sup>3-</sup> NC (Ag<sub>29</sub>-AC; AC denotes all-aroundcarboxyl functionalization; BDTA: 3,5-dithiolbenzoic acid; DPBA: 4-(diphenylphosphino)benzoic acid) by replacing the unfunctionalized carbon-tail-containing dithiolate and phosphine ligands of  $[Ag_{29}(BDT)_{12}(PPh_3)_4]^{3-}$  NC  $(Ag_{29}-BDT; BDT =$ 1,3-benzenedithiolate). The water solubility of Ag<sub>29</sub>-AC, while

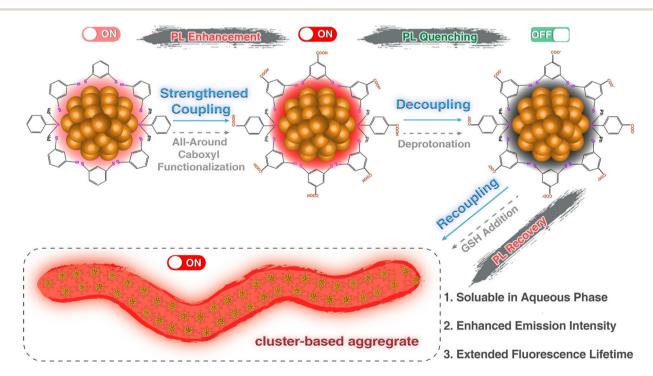


Fig. 5 Schematic of the emission enhancement of the Ag<sub>29</sub>-AC NC through the strengthened coupling, the emission quenching of Ag<sub>29</sub>-AC in the aqueous phase caused by molecular decoupling, and the emission recovery of the Aq<sub>29</sub>-AC NC by its supramolecular recoupling in the presence of reduced glutathione (GSH). Reproduced with permission from ref. 76. Copyright (2024).

retaining its size and structure, allows it to serve as a model NC compound to elucidate the effect of functionalization on PL. The quenching of PL (PLQY: 0.39%; lifetime: 0.21 µs) in Ag<sub>29</sub>-AC(H<sub>2</sub>O) clusters is triggered by phase transfer, leading to molecular decoupling. This increased mobility of Ag<sub>29</sub>-AC (H<sub>2</sub>O) clusters in water weakens the radiative transition. However, the emission recovery of the quenched NCs is achieved (PLOY: 4.7%; lifetime: 2.85 µs) through a supramolecular recoupling route induced by glutathione addition, which promotes the aggregation of NCs. This aggregation restricts intracluster motion and intercluster rotation, thereby strengthening the radiative transition of the clusters.

modification The selective peripheral  $[Ag_{29}(BDT)_{12}(TPP)_4]^{3-}$  NC with Ag(I) complexes, resulting in Ag<sub>29</sub> + Ag NCs, causes the PL peak to shift from 680 to 770 nm, accompanied by an increase in PLQY from 1.4 to 39%.77 Despite minor changes in the ground-state geometric and electronic structure, the excited-state properties undergo significant alterations. In addition to the inherent ultrafast intersystem crossing from the S<sub>1</sub> to T<sub>1</sub> state, the Ag<sub>29</sub> + Ag NC exhibits an additional excited-state relaxation to another triplet state (T'<sub>1</sub>) with a time constant of 287 ps. This additional relaxation may be associated with internal motions within the Ag<sub>29</sub> + Ag NC, such as rearrangement of the ion-pair structure. Such modifications hold promise for applications in photocatalysis.

Ligands and their functional groups modulate the structure and optical absorption features of NCs. The Ag<sub>14</sub> cluster structure is found to be robust, as it can be formed by a range of thiolate ligands, including monothiol and dithiol.<sup>38</sup> The dithiolated Ag<sub>14</sub> cluster exhibits unique absorption peaks at 425, 600, and 860 nm, extending into the near-infrared (NIR) region, primarily due to transitions derived from ligands. It also shows dual visible/NIR emission at approximately 680 and 997 nm. Conversely, the monothiolated Ag<sub>14</sub> cluster has visible absorption peaks at 368 and 530 nm, with yellow PL. 78 Due to the rigidity of the dithiolate, the resulting cluster is thermally robust. Furthermore, the electronic effects of halogen substituents on the thiolate ligand influence the optical and electrochemical bandgap values, governed by intracluster and intercluster  $\pi$ - $\pi$  interactions.<sup>79</sup> Additionally, the electronegativity of ligand substituents affects ultrafast relaxation dynamics by tuning the electron-phonon interaction.80 Intracluster interactions among thiolate and diphosphine ligands of Ag<sub>21</sub> clusters have been shown to enhance the PLQY up to 7%.81

In addition to ligands, electron-vibration coupling plays a crucial role in the PL of metal NCs. Weaker and stronger electron-vibration coupling leads to higher and lower NIR PLQY in  $Ag_{25}(SR)_{18}$  (3.5%) and  $Au_{25}(SR)_{18}$  (1%) clusters, respectively.<sup>82</sup> Doping a single Au atom into Ag<sub>25</sub> further suppresses the nonradiative decay rate, enhancing PLQY up to 35%.83 Pradeep and co-workers<sup>84</sup> improved the PL of Ag<sub>14</sub> NCs co-protected with naphthalenethiol and 1,6-bis(diphenylphosphino)hexane ligands by three-fold through secondary ligand-induced orthogonal self-assembly. The resulting supracolloidal structures, including helical fibers, spheres, and nanosheets, possess restricted intramolecular motion through noncovalent and supramolecular interactions, directly influencing radiative excited-state relaxation. Chiral NCs with light emission exhibit circularly polarized luminescence (CPL). The Zang group synthesized enantiomeric superatomic Ag clusters, 85 [Ag<sub>17</sub>(R/  $S-NYA)_{12}^{3+}$ , stabilized by chiral alkynyl ligands (R/S-NYA: N-((R/S)-1-(naphthalen-4-vl)ethyl)prop-2-vn-1-amine). exhibit NIR PLQY of 8% as well as CPL due to the chirality of the excited states.

# Transformation and surface reactivity

The molecular nature and intermediate stability of NCs render them reactive when subjected to external stimuli such as light, ligands, metal ions, temperature, and others. This reactivity often leads to their conversion into NCs of different sizes, including reduction or growth in size. Thus formed clusters may have novel electronic and geometric structures with exciting optoelectronic and photophysical properties.

#### 4.1. Light-induced transformation

colleagues<sup>86</sup> Pradeep reported [Ag<sub>42</sub>(CBDT)<sub>15</sub>(TPP)<sub>4</sub>]<sup>2-</sup> NC undergoes light-activated downsize conversion to form the Ag<sub>14</sub>(CBDT)<sub>6</sub>(TPP)<sub>6</sub> cluster (CBDT: ortho-carborane-1,2-dithiolate; TPP: triphenylphosphine). Despite the high thermal stability of Ag<sub>42</sub> cluster up to 200 °C in its solid-state, it undergoes conversion in solution under light irradiation (Fig. 6A). The transformation proceeds through two sets of intermediates (Fig. 6B): (i)  $[Ag_{37}(CBDT)_{12}(TPP)_4]^{3-}$  and  $[Ag_{35}(CBDT)_8(TPP)_4]^{2-}$  formed after 8 hours, and (ii)  $[Ag_{30}(CBDT)_8(TPP)_4]^{2-}$  $[Ag_{26}(CBDT)_{11}(TPP)_4]^{2-}$  $[Ag_{26}(CBDT)_7(TPP)_7]^{2-}$  formed after 16 hours. After 24 hours, these intermediates transform into the final product, the Ag<sub>14</sub> cluster. The kernel-centered stable excited states of Ag<sub>42</sub> are found to be responsible for light-induced cluster conversion. The Ag<sub>42</sub> cluster emits light at 980 nm with a phosphorescence lifetime of >1.5 μs, whereas Ag<sub>14</sub> shows red fluorescence emission at 626 nm with a lifetime of 550 ps.

In a recent study, the same group observed a light-activated size-expansion process, wherein the  $[Ag_{31}(TRZ)_{10}]^{2-/1-}$  cluster transformed into the [Ag<sub>42</sub>(TRZ)<sub>13</sub>]<sup>2+</sup> cluster (TRZ: 6-(dibutylamino)-1,3,5-triazine-2,4-dithiolate).87 The size conversion was demonstrated using absorption spectroscopy of NCs with fingerprint absorption features, as well as a clear color change from violet to greenish. The Ag<sub>31</sub> cluster exhibits photoresponsive behavior, and the resulting intermediate species are covered with solvated electrons, contributing to the expansion of the cluster size. Both the Ag<sub>31</sub> and Ag<sub>42</sub> clusters exhibit NIR and red emission, respectively. Notably, the final product Ag<sub>42</sub> cluster exhibits strikingly short excited-state carrier dynamics compared to the Ag<sub>31</sub> cluster. In other words, the stable photoexcited-state carriers and unique electronic and geometric structure of the Ag31 cluster likely lead to the light-induced size growth of NCs.

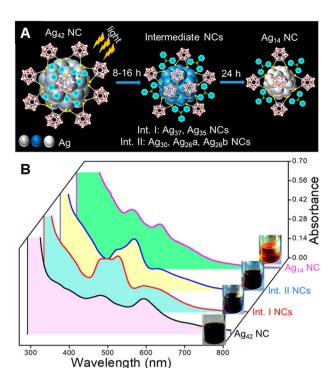


Fig. 6 (A) Schematic of the light-activated conversion of  $Ag_{42}$  to  $Ag_{14}$  through intermediate clusters. (B) UV-vis absorption spectra of  $Ag_{42}$  NC before and after light irradiation along with the mixture of intermediate clusters. Insets: photographs of the color change of cluster solutions during the light-irradiation. Reproduced with permission from ref. 86. Copyright (2021).

## 4.2. Ligand-exchange-induced transformation

When introduced to a foreign ligand, NCs undergo transformations depending on the inherent electronic effects of the new ligand and the stability of ligand-exchange intermediates. The Zhu group<sup>88</sup> utilized aqueous  $Ag_m(SG)_n$  clusters (where SG =L-glutathione) in a biphasic ligand-exchange process with two different thiols (HS<sup>t</sup>Bu: tert-butyl mercaptan; CySH: cyclohexanethiol). This process yielded Ag<sub>42</sub>(S<sup>t</sup>Bu)<sub>24</sub> and Ag<sub>61</sub>(SCy)<sub>40</sub>Cl. Ag<sub>42</sub> features a tetrahedral kernel, while Ag<sub>61</sub> possesses a facefused bi-tetrahedral Ag14 kernel. Interestingly, Ag42 can also convert to Ag<sub>61</sub> through a ligand etching process with excess HSCy. When Ag<sub>25</sub>(Capt)<sub>18</sub> NC (Capt: captopril) reacts with 2-isopropylbenzenethiol, the cluster size remains intact, 89 indicating that ligands determine both the cluster size and metal arrangements (i.e., structure). The  $[Ag_{18}H_{16}(PPh_3)_{10}]^{2+}$  clusters react with 2-pyrene imine thiol (2-PIT), forming the [Ag<sub>35</sub>(2-PIT)<sub>7</sub>(TPP)<sub>7</sub>@(H<sub>2</sub>O)]<sup>3+</sup> cluster.<sup>90</sup> This cluster emits blue and NIR light under UV excitation due to pyrene-to-metal core charge transfer and charge transfer within the metal core, respectively.

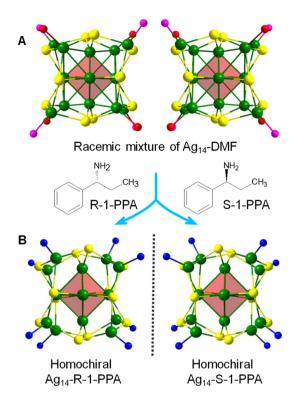
Liu and colleagues <sup>91</sup> observed the formation of a two-electron silver superatom,  $Ag_6(S_2P(O^iPr)_2)_4(dppm)_2$ , by reacting an eight-electron superatom,  $Ag_{20}(S_2P(O^iPr)_2)_{12}$ , with a diphosphine (dppm: bis(diphenylphosphino)methane). Here, dppm acts as a chemical scissor, pruning the icosahedron-based  $Ag_{20}$ 

NC to an octahedral Ag<sub>6</sub> NC. The Ag<sub>6</sub> cluster exhibits a bright yellow emission (PLQY: 16.3%) under UV light at room temperature. Sometimes, ligands react with clusters to form even smaller clusters or metal-ligand complex precursors. 92 When Pt<sub>2</sub>Ag<sub>23</sub>Cl<sub>7</sub>(PPh<sub>3</sub>)<sub>10</sub> <sup>93</sup> NCs react with a diphosphine ligand dppe (1,2-bis(diphenylphosphino)ethane), a self-assembled Ag<sub>2</sub>Cl<sub>2</sub>(dppe)<sub>2</sub> cluster material is obtained.<sup>94</sup> This material exhibits a high PLOY of ~18% due to metallophilic interactions and crystallization, which restrict intermolecular rotations and minimize nonradiative relaxation paths of the excited-state. Bidentate ligands can assemble the NCs if they have uncoordinated metal sites. The 4,4'-bipyridine (4,4'-bpy) links two twoelectron primary building units Ag<sub>10</sub>(S<sub>2</sub>P(O<sup>i</sup>Pr)<sub>2</sub>)<sub>8</sub> to make the smallest self-assembled structure of [Ag<sub>10</sub>(S<sub>2</sub>P(O<sup>i</sup>Pr)<sub>2</sub>)<sub>8</sub>]<sub>2</sub>(μ-4,4'bpy)<sup>95</sup> through N coordination with open Ag site of the cluster. Upon heat treatment, an Ag(1) hydride cluster Ag<sub>7</sub>(H)(S<sub>2</sub>P (O<sup>i</sup>Pr)<sub>2</sub>)<sub>6</sub> undergoes self-redox reaction to produce Ag<sub>10</sub>(S<sub>2</sub>P (O<sup>i</sup>Pr)<sub>2</sub>)<sub>8</sub> cluster without addition of any reducing agent. 96 The bicapped trigonal bipyramid metal framework of Ag<sub>7</sub> cluster is well correlated with that of Ag<sub>10</sub>, suggesting Ag<sub>7</sub> acts as the template in the growth of the cluster.

#### 4.3. Surface reactivity

When a cluster possesses accessible surfaces or labile ligands with bulkiness, it tends to exhibit distinct reactivity, which can be harnessed for specific applications through materials design. For example, the  $[Ag_{29}(SSR)_{12}(PPh_3)_4]^{3-}$  cluster is hydrophobic, limiting its utility in applications involving aqueous media. However, the accessible surface and high magnitude of negative electronic charge of this robust cluster enable surface interactions with solvent-conjoined Na<sup>+</sup> cations  $(Na_1(NMP)_5 \text{ or } Na_3(DMF)_{12}) (NMP: N-Methyl-2-pyrrolidone).$ The strong electrostatic attraction between the anionic cluster and these positive ions facilitates the migration of the clusters  $[Ag_{29}(SSR)_{12}(PPh_3)_4]^{3}$   $[Na_1(NMP)_5]_3$ aqueous medium, thus enabling biological applications. Such a micellization strategy has been found to be applicable to most anionic clusters.

Zheng and colleagues98 synthesized a solvent-coordinated chiral Ag cluster,  $[Ag_{14}(SPh(CF_3)_2)_{12}(PPh_3)_4(DMF)_4]$   $(Ag_{14}-DMF)_6$ in which surface ligands are highly labile. The high surface reactivity of Ag<sub>14</sub>-DMF allows DMF to dissociate or be exchanged with new ligands. Utilizing the surface reactivity of the cluster, a fully amine-ligated Ag14 cluster (where all PPh3 and DMF are replaced) could be obtained with fine-tuned optical properties by exposing the cluster to excess amine. Interestingly, by introducing chiral amines (R-1-PPA: R-1phenylpropylamine and S-1-PPA: S-1-phenylpropylamine), the racemic Ag<sub>14</sub>-DMF clusters (Fig. 7A) are converted to homochiral Ag<sub>14</sub> clusters (Fig. 7B) with a good chiroptical response. By employing bidentate linkers with different chiral configurations, <sup>99</sup> such as  $(1R,2R,N^1E,N^2E)-N^1,N^2$ -bis(pyridin-3-ylmethylene)cyclohexane-1,2-diamine (LR) and its chiral analog LS, chiral cluster-based metal-organic frameworks are synthesized using Ag14-DMF clusters. These cluster-organic frameworks demonstrate assembly-disassembly process.



**Fig. 7** Core structures of (A) Ag<sub>14</sub>-DMF enantiomer pair and (B) Ag<sub>14</sub>-R-1-PPA (left) and Ag<sub>14</sub>-S-1-PPA (right). Reproduced with permission from ref. 98. Copyright (2021).

The surface reactivity of NCs with metal ions can also serve as a useful strategy to obtain novel NCs. Zhu and co-workers 100 reported on the role of metal ion quantity in NC transformation. When [Ag<sub>25</sub>(2,5-DMBT)<sub>16</sub>(DPPF)<sub>3</sub>]<sup>+</sup> reacts with less than 0.5 equiv. of  $Cu^{2+}$ , size-retained doped clusters  $[Ag_{25-x}Cu_x(2,5-x)]$ DMBT)<sub>16</sub> $(DPPF)_3$ ]<sup>+</sup> are formed (2,5-DMBT: 2,5-dimethylbenzenethiolate; DPPF: 1,1'-bis(diphenylphosphino)ferrocene). On the other hand, one equiv. of Cu<sup>2+</sup> results in a structural process, transformation yielding  $[Ag_{22-x}Cu_{x}(2,5-$ DMBT)<sub>12</sub>(DPPF)<sub>4</sub>Cl<sub>4</sub>]<sup>2+</sup> clusters. In the former case, Cu<sup>2+</sup> acts as a dopant, while in the latter case, it acts as an oxidant. Control experiments demonstrate that Cu<sup>+</sup> forms doped clusters, whereas Cu<sup>2+</sup> and other common oxidizers produce sizetransformed products, indicating the unique role of metal ions in NC reactions.

# 5. Catalytic applications

Metal NCs with well-defined size, structure, and composition can serve as model nanoparticle analogs to understand the molecular-level mechanisms of catalytic properties. In the following, we will cover various reactions catalyzed by Ag NCs.

#### 5.1. CO<sub>2</sub> conversion

In addressing global energy demands and closing the carbon cycle, CO<sub>2</sub> conversion plays a crucial role. Electrocatalysis

holds great potential for converting  $CO_2$  into useful fuels and chemicals using electrical energy produced from renewable sources. Recently, Ag NCs have begun to be utilized for  $CO_2$  reduction. Tang and colleagues<sup>101</sup> synthesized a body-centered cubic (bcc) structured homoleptic  $Ag_{15}(C \equiv C^{-t}Bu)_{12}^{+}$  cluster and employed it as a catalyst for the electrochemical  $CO_2$  reduction reaction (eCO<sub>2</sub>RR). The cluster catalyst exhibits high selectivity for CO, with a highest Faradaic efficiency (FE<sub>CO</sub>) of 95% at -0.6 V and a maximal turnover frequency of 6.37 s<sup>-1</sup> at -1.1 V, along with excellent long-term stability. Theoretical calculations indicate that the Ag site created upon stripping a single alkynyl ligand acts as the active site.

In another study, the same group utilized two Ag NCs of identical size, but one was protected by alkynyl ligands and the other was co-protected by thiolate and diphosphine ligands to understand the effect of ligands on eCO<sub>2</sub>RR. <sup>102</sup> The clusters are Ag<sub>32</sub>L<sub>24</sub> (L = 3,5-bis(trifluoromethylbenzene)acetyand  $[Ag_{32}(dppe)_5(SR)_{24}]^{2-}$ (SR: trifluoromethylbenzenethiolate). The alkynyl-protected Ag<sub>32</sub> cluster exhibits a high FE<sub>CO</sub> of 96.4% at -0.8 V, whereas the thiolated cluster shows an FE<sub>CO</sub> of only 56.6% at -1.0 V, suggesting the critical role of ligands in eCO2RR. Theoretical calculations indicate that the energy barrier for the formation of the \*COOH intermediate is similar for both clusters. However, the thiolated cluster has the lowest thermodynamic barrier for the formation of H<sub>2</sub> gas (0.05 eV) compared to that (0.51 eV) of the alkynylated cluster, explaining the high CO selectivity in the latter case.

The thiacalix[4]arene-capped 6-electron [Ag<sub>30</sub>(TC4A)<sub>4</sub>(<sup>i</sup>PrS)<sub>8</sub>] NC and the chain-like Ag(i) polymer [H<sub>2</sub>Ag<sub>5</sub>(TC4A)(<sup>i</sup>PrS)<sub>3</sub>] $\infty$  are compared with their eCO<sub>2</sub>RR performance. <sup>103</sup> The NC exhibits high CO selectivity (93.4% at -0.9 V *versus* RHE) and long-term stability (24 h). Density functional theory (DFT) calculations show that the stabilization of the \*COOH intermediate is better on the cluster than on the polymer. Furthermore, the  $j_{\rm CO}$  of Ag<sub>30</sub> is -33.29 mA cm<sup>-2</sup> at -1.2 V *versus* RHE, which is 2.5 times higher than that of the Ag(i) polymer (-13.2 mA cm<sup>-2</sup>).

Recently, we have demonstrated that the local hydrophobicity of the surface of NCs influences eCO2RR by affecting the interfacial water structure.89 To understand the molecular-level mechanism, two Ag NCs, namely Ag<sub>25</sub>(Capt)<sub>18</sub> and Ag<sub>25</sub>(IPBT)<sub>18</sub>, were synthesized. They exhibit identical metalligand arrangements except for the difference in the ligand structure, wherein the carboxyl group of the Capt ligand makes the NC water-soluble, while the isopropyl on the benzene ring of the IPBT ligand makes the NC hydrophobic. Despite loading the NCs on a carbon support, the local hydrophilicity of Ag<sub>25</sub>(Capt)<sub>18</sub> is confirmed using contact angle measurements. The eCO<sub>2</sub>RR activity of the Ag<sub>25</sub>(IPBT)<sub>18</sub> cluster is dramatically enhanced by locally induced hydrophobicity due to bulky alkyl functionality near its surface (Fig. 8A). This enhancement is reflected in the FE<sub>CO</sub> (90.3%) and higher  $j_{CO}$ observed in an H-cell compared to the Ag<sub>25</sub>(Capt)<sub>18</sub> cluster with confined hydrophilicity (FE<sub>CO</sub>: 66.6%) (Fig. 8B). Notably, the hydrophobic  $Ag_{25}$  cluster exhibits an impressive  $j_{CO}$  as

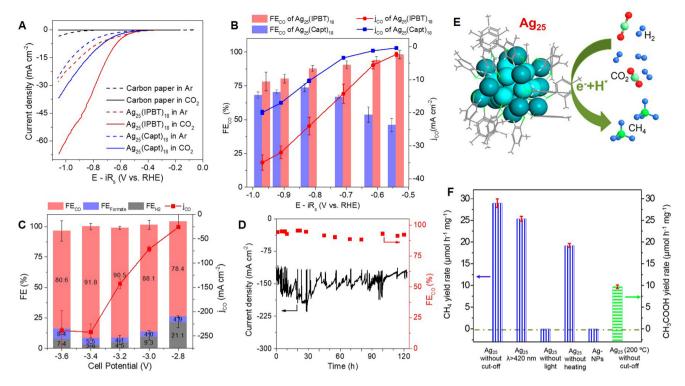


Fig. 8 (A) Polarization curves and (B)  $FE_{CO}$  and  $j_{CO}$  of  $Ag_{25}(IPBT)_{18}$  and  $Ag_{25}(Capt)_{18}$  catalysts in the H-cell. (C)  $FE_{CO}$  and  $j_{CO}$  of  $Ag_{25}(IPBT)_{18}$  in the cell potential range from -2.8 to -3.6 V in the MEA cell. (D) Long-term operation of  $Ag_{25}(IPBT)_{18}$  at -3.2 V cell potential in the MEA cell. Reproduced with permission from ref. 89. Copyright (2024). (E) Structure and photocatalytic  $CO_2$  hydrogenation of  $Ag_{25}(IPBT)_{18}$  cluster. (F)  $CH_4$  and  $CH_3COOH_5$  yields obtained via  $CO_2$  hydrogenation over  $Ag_{25}$  clusters ( $CO_2: H_2 1: 3 \text{ v/v}$ , 1.5 MPa). Reproduced with permission from ref. 104. Copyright (2021).

high as  $-240~\text{mA}~\text{cm}^{-2}$  with  $\text{FE}_{\text{CO}} > 90\%$  at a cell potential of -3.4~V in a gas-fed membrane electrode assembly device (Fig. 8C). Furthermore, this cluster demonstrates stable  $\text{eCO}_2\text{RR}$  activity over a period of 120 hours (Fig. 8D). Operando surface-enhanced infrared absorption spectroscopy and theoretical simulations reveal that the hydrophilic cluster surface accommodates more water molecules due to hydrogen bonding. In contrast, the interfacial water structure on the hydrophobic cluster is similar to bulk water. These differences influence the thermodynamics and kinetics of  $\text{eCO}_2\text{RR}$ .

Jin and co-workers 104 utilized the Ag<sub>25</sub>(DMBT)<sub>18</sub> cluster (DMBT: 2,4-dimethylbenznethiolate) for photocatalytic CO<sub>2</sub> hydrogenation (Fig. 8E). The product selectivity is 100% towards CH<sub>4</sub> at a relatively mild temperature (100 °C). CO<sub>2</sub> adsorption on Ag<sub>25</sub> is found to be energetically favorable. Operando infrared spectroscopy reveals CO2 methanation proceeds through a -H-assisted multielectron reaction pathway via the transformation of formyl and formaldehyde species into surface CH<sub>x</sub>. CO<sub>2</sub> photoreduction is promoted by Ag<sub>25</sub> clusters, yielding methane at a rate of 28.95 µmol h<sup>-1</sup> mg<sup>-1</sup> at 100 °C under Xe lamp illumination without a cutoff filter (Fig. 8F). With a 420 nm cutoff filter, Ag<sub>25</sub> produces CH<sub>4</sub> at a rate of 25.34 µmol h<sup>-1</sup> mg<sup>-1</sup>. No activity is observed without light, and heating promotes CO2 photoreduction rate. Interestingly, the selectivity switches from CH<sub>4</sub> to acetic acid with a rise in bed temperature. It is suggested that the CO intermediate might bind with Ag-CH<sub>3</sub> intermediate at high

temperatures (200 °C) to form acetic acid. The Ag NCs stabilized by lacunary polyoxometalates,  $[Ag_{24}(Si_2W_{18}O_{66})_3]$  and  $[Ag_{27}(Si_2W_{18}O_{66})_3]$ , form HCOOH with 90% selectivity, along with minor products CO and  $H_2.^{105}$  This result indicates that tuning product selectivity may be possible by changing the ligand shell of the NCs.

#### 5.2. Hydrogen evolution

The Pinna group deposited  $Ag_{44}(SR)_{30}$  NCs onto a wide-bandgap semiconductor  $TiO_2$ . It produces hydrogen under simulated solar light from water splitting at a rate of 7.4 mmol  $h^{-1}$   $g_{catalyst}^{-1}$ , which is three orders of magnitude higher than under visible light irradiation. The former value is five times higher than that of  $TiO_2$  modified with plasmonic Ag nanoparticles, indicating the importance of NCs in photocatalysis. Energy band alignment and transient absorption spectroscopy reveal the formation of a type II heterojunction charge-transfer route under UV-vis irradiation, with the cluster acting as a small bandgap semiconductor. Consequently, clusters function as cocatalysts rather than mere photosensitizers.

Synthesizing core–shell nanomaterials with atomic-level control poses significant challenges despite their potential for enhancing optical and catalytic properties. We reported the synthesis and crystal structure of  $[Au_{12}Ag_{32}(SePh)_{30}]^{4-}$ , the first instance of selenolated Au–Ag core–shell NCs.<sup>107</sup> These clusters feature a gold icosahedron core encased within a silver dodecahedron, protected by an  $Ag_{12}(SePh)_{30}$  shell. The pres-

ence of the gold core strongly influences the overall electronic structure, leading to synergistic effects that enhance stability and near-infrared-II photoluminescence. Both Au<sub>12</sub>Ag<sub>32</sub> and its homometal analog Ag44 exhibit strong interactions with oxygen vacancies on TiO2, facilitating interfacial charge transfer for photocatalysis. Indeed, Au<sub>12</sub>Ag<sub>32</sub>/TiO<sub>2</sub> demonstrates remarkable solar  $H_2$  production (6810  $\mu$ mol  $g^{-1}$   $h^{-1}$ ), surpassing Ag<sub>44</sub>/TiO<sub>2</sub> and TiO<sub>2</sub> by approximately 6.2 and 37.8 times, respectively (Fig. 9A and B). Furthermore, Au<sub>12</sub>Ag<sub>32</sub>/TiO<sub>2</sub> exhibits good stability and recyclability with minimal catalytic activity loss (Fig. 9C and D). Experimental and computational findings suggest that Au<sub>12</sub>Ag<sub>32</sub> serves as an efficient cocatalyst due to its favorable electronic structure, which aligns well with TiO<sub>2</sub> bands, promoting enhanced separation of photoinduced charge carriers, facilitated by the relatively negatively charged Au<sub>12</sub> core.

Heteroatom doping may not always have a positive effect on photocatalysis. For instance, when the Ag<sub>25</sub> cluster is doped with Au and Pd to form isostructural centrally doped AuAg<sub>24</sub> and PdAg24 clusters, respectively, the photocatalytic hydrogen production is lower than that of the parent Ag<sub>25</sub> cluster. 108 Furthermore, the hydrogen production rate of Ag<sub>25</sub>/TiO<sub>2</sub> is 10 times higher than that of TiO2 alone due to the enhanced charge separation at the interface. Measurements using ultraviolet photoelectron spectroscopy and DFT calculations indicate that the lower activities of doped clusters result from an energy mismatch between the levels of doped NCs and TiO<sub>2</sub>.

Zhu and co-workers<sup>109</sup> systematically designed an electrocatalytic hydrogen evolution catalyst based on an Ag<sub>29</sub> cluster by core-alloying with Pt and grafting optimal Mn onto the cluster's surface. In the resulting catalyst, PtAg28-BTT-Mn(10), the Mn site exhibits a  $\Delta G_H^*$  value of 0.18 eV, which is suitable for the hydrogen evolution reaction. This suggests that the Mn sites, rather than the Ag<sub>3</sub> faces of the Ag<sub>29</sub> cluster, serve as the catalytic sites.

#### 5.3. Nitrate reduction

The Zang group<sup>110</sup> loaded [Ag<sub>9</sub>(mba)<sub>9</sub>](NH<sub>4</sub>)<sub>9</sub> NCs onto Ti<sub>3</sub>C<sub>2</sub> MXene to form Ag<sub>9</sub>/MXene composite catalyst for highly efficient nitrate reduction to ammonia under ambient con-

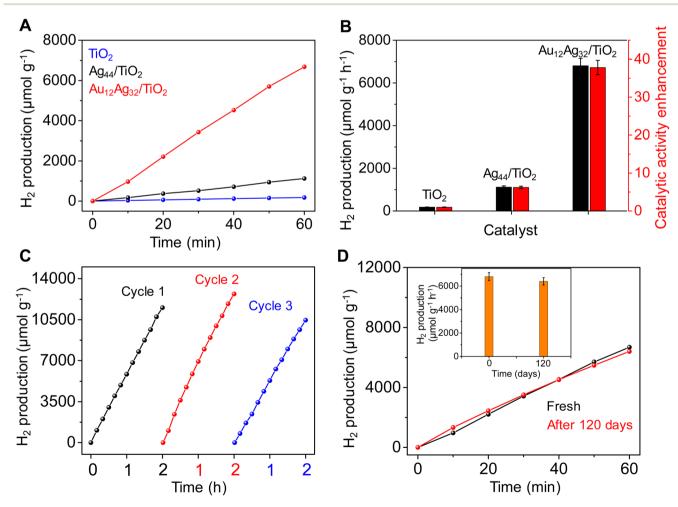


Fig. 9 (A and B) Photocatalytic H<sub>2</sub> production and improvement with Ag<sub>44</sub>/TiO<sub>2</sub> and Au<sub>12</sub>Ag<sub>32</sub>/TiO<sub>2</sub> compared to TiO<sub>2</sub>. (C) H<sub>2</sub> evolution activity of recycled Au<sub>12</sub>Ag<sub>32</sub>/TiO<sub>2</sub>. (D) H<sub>2</sub> production by fresh and aged Au<sub>12</sub>Ag<sub>32</sub>/TiO<sub>2</sub>. Inset: Bar diagram demonstrating nearly unchanged photocatalytic activity of Au<sub>12</sub>Ag<sub>32</sub>/TiO<sub>2</sub> clusters after four months of storage. Reproduced with permission from ref. 107. Copyright (2023).

ditions with improved stability in neutral medium. The Ag clusters promote the conversion of nitrate to nitrite. The nitrite-enriched MXene promotes the sustained reaction of nitrite conversion to ammonium ion, establishing a tandem catalytic system. Furthermore, compared to solely Ag<sub>9</sub> NCs, Ag<sub>9</sub>/MXene shows enhanced catalytic stability, with no decay in current density after 108 hours of reaction. Tang and colleagues<sup>111</sup> synthesized a homoleptic alkynyl-protected twoelectron  $\left[Ag_{30}Pd_4(C_6H_9)_{26}\right]^{2+}$  cluster and employed it for nitrate to ammonia reduction. The Pd atoms occupy the subcenter of the metal core. This Ag<sub>30</sub>Pd<sub>4</sub> cluster acts as an efficient electrocatalyst for ammonia production with a selectivity of over 90%. In situ infrared spectroscopy results show that nitrate to nitrite formation takes place at the silver site, and Pd sites contribute majorly to the nitrite to ammonia conversion. These experimental observations are consistent with DFT calculations, wherein nitrate preferably binds with silver. Upon binding with a water molecule, nitrite is released. This nitrite is then transferred to the vicinally exposed Pd site to complete the reaction. These results demonstrate tandem catalysis of NCs rather than the usually thought synergistic catalysis.

#### 5.4. Other energy related applications

A molecular hybrid of an atomically precise silver NC and polyoxometalates is found to cleave H2 into protons and electrons, which are stored at the polyoxometalate surface and silver, respectively. 112 The cluster  $(Ag_{27})^{17+}$  accommodates electrons by becoming  $(Ag_{27})^{13+}$ . The polyoxometalate stabilizes the unstable Ag cluster without structural transformation, providing a unique interface for H<sub>2</sub> dissociation and storing protons on its negatively charged basic surface. This example may open new avenues for fuel cell applications. A stable supercapacitor electrode is fabricated using a two-dimensional highnucleus Ag NC. 113 The high stability of the Ag NC polymer, enabled by its strong coordination with decafluoroazelaic acid (CF2), exhibits a high energy density of 372 F  $g^{-1}$  at 4.5 A  $g^{-1}$ , along with excellent cycling stability and capacity. This NC demonstrates both high power density and long cycle life, enduring over 6000 cycles with a retention rate of 95%. Additionally, it can tolerate a wide range of scan rates from 5 mV s<sup>-1</sup> to 1 V s<sup>-1</sup>, suggesting its potential as a sustainable energy source.

Ag NCs catalyse the hypergolic activity of inert o-carboranealkynyl ligand when they contact with white fuming nitric acid oxidizer. 114 Spontaneous ignition and combustion produce heat, making o-carboranealkynyl-protected Ag clusters as potential components of propellant systems. Similarly, Ag cluster-based organic frameworks also exhibit high-performance hypergolic properties. 115 The graphitic carbon nitride  $(g-C_3N_4)$ -loaded  $Ag_{33}(4-MePhC \equiv C)_{22}(Dppp)_4$  NC serves as excellent photocatalyst for degradation of methyl orange. 116 Compared to pristine g-C<sub>3</sub>N<sub>4</sub>, the cluster-loaded catalyst shows a 16-fold activity enhancement due to the enhanced separation of photogenerated charge carriers. The co-assembled  $[Ag_{29}(BDT)_{12}(TPP)_2]^{3-}$  clusters and  $Ru(bpy)_3^{2+}$  photosensitizer (i.e., Ag<sub>29</sub>Ru coassembly) exhibits a greatly enhanced singlet oxygen production capacity under visible light irradiation as compared to pristine Ag<sub>29</sub> cluster. 117 Narrowing of the HOMO-LUMO gap of the Ag<sub>29</sub>Ru coassembly is suggested to be the reason for enhancement in singlet oxygen production.

### 5.5. Organic transformation

The Wang group<sup>118</sup> synthesized two identically sized Ag NCs with different types of ligands, enabling the study of the effect of ligands on the stability and catalytic activity trade-off in the nitro compounds. The hydrogenation of clusters  $\left[Ag_{21}(H_2BTCA)_3(O_2PPh_2)_6\right]^+$ and  $[Ag_{21}(C = CC_6H_3 - 3, 5 R_2$ <sub>06</sub> $(O_2PPh_2)_{10}$ <sup>+</sup> (Fig. 10A and B, respectively) where  $H_4BTCA$ represents p-tert-butylthiacalix[4]arene and R represents OMe, possess identical icosahedral kernels. The thiacalix[4]areneprotected Ag<sub>21</sub> is found to be an 8-electron superatomic cluster, exhibiting robust stability under high temperatures (90 °C) in solution and oxidative conditions (30% H<sub>2</sub>O<sub>2</sub>). On the other hand, the alkynylated Ag<sub>21</sub> is a four-electron cluster with an open-shell electronic configuration, demonstrating good ambient stability. However, the alkynylated cluster could catalyse nitrophenol reduction to aminophenol (Fig. 10C) and complete the reaction in 16 minutes (rate constant: 0.228 min<sup>-1</sup>), while the other cluster could not finish the reaction even after 1.5 hours (rate constant: 0.023 min<sup>-1</sup>). Notably, the thiacalix[4]arene-protected Ag<sub>21</sub> shows significantly higher catalytic activity than the catalyst support. These results highlight the importance of the electronic structure of clusters in catalysis, which can be modulated by an appropriate combination of protecting ligands.

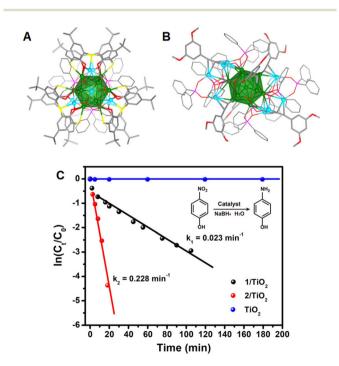


Fig. 10 Molecular structures of (A) thiacalixarene- and (B) alkynyl-protected Ag<sub>21</sub> NCs. (C) Catalytic activity of Ag<sub>21</sub>(thiacalixarene)/TiO<sub>2</sub> (1/ TiO<sub>2</sub>) and Ag<sub>21</sub>(alkynyl)/TiO<sub>2</sub> (2/TiO<sub>2</sub>) and TiO<sub>2</sub> in nitrophenol reduction. Reproduced with permission from ref. 118. Copyright (2022).

group<sup>119</sup> The Chen synthesized three identical Ag<sub>33</sub>(SR)<sub>24</sub>(PPh<sub>3</sub>)<sub>4</sub> NCs with three different thiolate ligands: 2-phenylethanethiol, 4-chlorobenzylmercaptan, and 4-methoxybenzylmercaptan. The benzylthiolated Ag<sub>33</sub> NCs exhibit superior hydrogenation activity compared to the phenylethanethiolated Ag<sub>33</sub> cluster. This ligand-dependent catalysis is attributed to the noncovalent  $(H-\pi$  and  $\pi-\pi)$  interactions of ligands and substrates through phenyl rings, facilitating facile adsorption on the cluster catalyst. Furthermore, the cluster catalysts can be recycled more than five times without significantly losing activity. The surface-exposed Ag<sub>30</sub> clusters are incorporated into molecular metal oxide cavities, 120 which demonstrate exceptional hydrogenation activities using H2 gas. The substrate scope is very broad for these cluster catalysts with high yields.

The Ag<sub>25</sub>(SR)<sub>18</sub> NCs serve as templates to fabricate AgPd bimetallic nanocatalysts with controlled Pd content and distribution. 121 Initially, the Ag<sub>25</sub>(SR)<sub>18</sub> NCs are activated at low temperature, acting as a template for subsequent Pd deposition. By employing low and high amounts of Pd, the location of Pd is manipulated at the subsurface and surface, respectively. The AgPd nanostructure with Pd at the subsurface (12:1-Ag:Pd/ carbon) exhibits controlled alkyne (2-methyl-3-butyn-2-ol) to alkene (2-methyl-3-buten-2-ol) hydrogenation with 96.6% selectivity. On the other hand, the catalyst with a high Pd content (1:6-Ag:Pd/carbon) shows fully hydrogenated products, suggesting the use of Ag NCs in catalyst design. Upon controlled activation of Ag<sub>25</sub> clusters at 250 °C with an unchanged cluster size, 122 they function as excellent catalysts for styrene oxidation with high selectivity for styrene oxide. However, above this temperature, the clusters sinter to form nanoparticles, and the catalytic performance decreases dramatically with poor styrene oxide selectivity. The Ag<sub>4</sub> NC is synthesized for catalyzing the cyclization of propargylamine with CO<sub>2</sub>. 123 The alkyne-captured cluster catalyst intermediate is confirmed by mass spectrometry. This cluster catalyst exhibits a broad substrate scope with high turnover number values and recyclability for several cycles. Furthermore, gram-scale cyclized products could be achieved using the Ag<sub>4</sub> cluster-based catalyst.

#### Summary and perspectives 6.

We have conducted a brief overview of the recent advancements in Ag(0)-containing atomically precise Ag nanoclusters, highlighting the innovative strategies employed in their design and synthesis. By leveraging a rare combination of ligands and meticulously crafting atomic-level structures, researchers have significant progress in enhancing stability, achieved functionalization, and tuning optical and photophysical properties of these nanoclusters. Moreover, the transformative potential of these clusters in catalytic applications, particularly in understanding the role of structures and ligands in catalytic activity and selectivity, has been extensively discussed.

Despite these strides, several challenges remain to be addressed to propel Ag nanocluster research forward. Chief among these is the need to synthesize a greater variety of nanoclusters with consistent size, atomic arrangement, and ligand environment, enabling a comprehensive understanding of their dependency on optical and photophysical characteristics. This necessitates the development of novel synthetic routes (e.g., solidstate 124,125 and biphasic 89 reactions) and the exploration of unconventional precursors, solvents, and reducing agents. Additionally, efforts to upscale synthesis for practical applications must be intensified, requiring advancements in design principles and synthetic chemistry to increase yield and scalability.

The intriguing transformation chemistry of Ag nanoclusters, especially under light-induced conditions, presents an exciting avenue for future exploration. By precisely controlling irradiation parameters and capturing intermediate clusters, researchers can unlock new catalytic possibilities. Furthermore, there is a pressing need to synthesize Ag nanoclusters analogous to existing Au counterparts, such as the elusive Ag<sub>38</sub>(SR)<sub>24</sub> analogue, to facilitate comparative studies and expand our understanding of chiroptical and catalytic properties.

While the applications of Ag nanoclusters have been explored to a limited extent, their lower stability compared to gold counterparts presents both challenges and opportunities. The metastability of Ag clusters can be leveraged to create novel catalysts by probing interactions with support materials and elucidating photo-excited-state dynamics. Additionally, the utilization of Ag clusters in CO<sub>2</sub> electrolyzers, owing to their high atom utilization efficiency, holds promise for practical applications. Exploring interactions with other clusters, metal ions, ligands, surfaces, and interfaces will undoubtedly unlock new avenues and possibilities in cluster chemistry. The use of heterometal atoms during the synthesis of silver nanoclusters may lead to either doped silver clusters or novel Ag clusters different from those obtained in the absence of heterometal precursors. Although predicting the choice of heterometal precursors is challenging, considering their ability to form complexes with ligands or transient alloy species may be helpful in assisting the controlled nucleation and growth stages. The heteroatom-doped silver clusters are anticipated to exhibit enhanced catalytic performance and stability due to synergistic and cooperative effects.

In conclusion, the future of cluster chemistry is exceptionally bright, with Ag nanoclusters poised to play a pivotal role in advancing catalysis, energy conversion, and materials science. Continued interdisciplinary efforts and innovative approaches will be essential in harnessing the full potential of these atomically precise nanomaterials.

## Conflicts of interest

There are no conflicts to declare.

# Acknowledgements

This work was financially supported by the Institute for Basic Science (IBS) with Project Codes IBS-R006-D1 and IBS-R006-Y2.

## References

- 1 I. Chakraborty and T. Pradeep, *Chem. Rev.*, 2017, **117**, 8208–8271.
- 2 Y. Du, H. Sheng, D. Astruc and M. Zhu, Chem. Rev., 2020, 120, 526–622.
- 3 R. Jin, G. Li, S. Sharma, Y. Li and X. Du, *Chem. Rev.*, 2021, 121, 567–648.
- 4 M. F. Matus and H. Häkkinen, Nat. Rev. Mater., 2023, 8, 372-389.
- 5 Y. Negishi, K. Nobusada and T. Tsukuda, *J. Am. Chem. Soc.*, 2005, **127**, 5261–5270.
- 6 Z. Wang, M.-D. Li, J.-Y. Shi, H.-F. Su, J.-W. Liu, L. Feng, Z.-Y. Gao, Q.-W. Xue, C.-H. Tung, D. Sun and L.-S. Zheng, CCS Chem., 2022, 4, 1788–1795.
- 7 Y.-M. Su, B.-Q. Ji, Z. Wang, S.-S. Zhang, L. Feng, Z.-Y. Gao, Y.-W. Li, C.-H. Tung, D. Sun and L.-S. Zheng, *Sci. China: Chem.*, 2021, 64, 1482–1486.
- 8 S.-S. Zhang, R.-C. Liu, X.-C. Zhang, L. Feng, Q.-W. Xue, Z.-Y. Gao, C.-H. Tung and D. Sun, *Sci. China: Chem.*, 2021, 64, 2118–2124.
- 9 Z. Wang, Q.-P. Qu, H.-F. Su, P. Huang, R. K. Gupta, Q.-Y. Liu, C.-H. Tung, D. Sun and L.-S. Zheng, *Sci. China: Chem.*, 2020, 63, 16–20.
- Z. Wang, H.-F. Su, G.-L. Zhuang, M. Kurmoo, C.-H. Tung,
   D. Sun and L.-S. Zheng, CCS Chem., 2020, 2, 663–672.
- 11 S. Takano and T. Tsukuda, J. Am. Chem. Soc., 2021, 143, 1683–1698.
- 12 K. Zheng, X. Yuan and J. Xie, *Chem. Commun.*, 2017, 53, 9697–9700.
- 13 H. Lin, X. Song, O. J. H. Chai, Q. Yao, H. Yang and J. Xie, Adv. Mater., 2024, DOI: 10.1002/adma.202401002.
- 14 X. Liu, J. Chen, J. Yuan, Y. Li, J. Li, S. Zhou, C. Yao, L. Liao, S. Zhuang, Y. Zhao, H. Deng, J. Yang and Z. Wu, Angew. Chem., Int. Ed., 2018, 57, 11273–11277.
- 15 N. Xia, J. Yang and Z. Wu, *Nanoscale*, 2015, 7, 10013–10020.
- 16 S. Biswas, A. K. Das and S. Mandal, Acc. Chem. Res., 2023, 56, 1838–1849.
- 17 G. Deng, K. Lee, H. Deng, M. S. Bootharaju, N. Zheng and T. Hyeon, *J. Phys. Chem. C*, 2022, **126**, 20577–20583.
- 18 K. Lee, G. Deng, M. S. Bootharaju and T. Hyeon, *Acc. Chem. Res.*, 2023, **56**, 1118–1127.
- 19 W. Jing, H. Shen, R. Qin, Q. Wu, K. Liu and N. Zheng, *Chem. Rev.*, 2023, **123**, 5948–6002.
- 20 T. Kawawaki, Y. Kataoka, M. Hirata, Y. Akinaga, R. Takahata, K. Wakamatsu, Y. Fujiki, M. Kataoka, S. Kikkawa, A. S. Alotabi, S. Hossain, D. J. Osborn, T. Teranishi, G. G. Andersson, G. F. Metha, S. Yamazoe and Y. Negishi, *Angew. Chem., Int. Ed.*, 2021, 60, 21340– 21350.
- 21 T. Kawawaki, T. Okada, D. Hirayama and Y. Negishi, *Green Chem.*, 2024, **26**, 122–163.
- 22 A. K. Das, S. Biswas, A. Pal, S. S. Manna, A. Sardar, P. K. Mondal, B. Sahoo, B. Pathak and S. Mandal, *Nanoscale*, 2024, **16**, 3583–3590.

- 23 Y. Yun, H. Sheng, K. Bao, L. Xu, Y. Zhang, D. Astruc and M. Zhu, J. Am. Chem. Soc., 2020, 142, 4126–4130.
- 24 S. Wang, S. Jin, S. Yang, S. Chen, Y. Song, J. Zhang and M. Zhu, *Sci. Adv.*, 2015, 1, e1500441.
- 25 A. Desireddy, B. E. Conn, J. Guo, B. Yoon, R. N. Barnett, B. M. Monahan, K. Kirschbaum, W. P. Griffith, R. L. Whetten, U. Landman and T. P. Bigioni, *Nature*, 2013, 501, 399–402.
- 26 H. Yang, Y. Wang, H. Huang, L. Gell, L. Lehtovaara, S. Malola, H. Häkkinen and N. Zheng, *Nat. Commun.*, 2013, 4, 2422.
- 27 H. Chang, M. S. Bootharaju, S. Lee, J. H. Kim, B. H. Kim and T. Hyeon, *Bull. Korean Chem. Soc.*, 2021, 42, 1386–1399.
- 28 X.-H. Ma, Y. Si, L.-L. Luo, Z.-Y. Wang, S.-Q. Zang and T. C. W. Mak, *ACS Nano*, 2022, **16**, 5507–5514.
- 29 X. Ma, Y. Bai, Y. Song, Q. Li, Y. Lv, H. Zhang, H. Yu and M. Zhu, *Angew. Chem., Int. Ed.*, 2020, 59, 17234–17238.
- 30 G. Deng, T. Ki, S. Yoo, X. Liu, K. Lee, M. S. Bootharaju and T. Hyeon, *Nanoscale*, 2024, DOI: 10.1039/D4NR01471E.
- 31 F. Hu, R.-L. He, Z.-J. Guan, C.-Y. Liu and Q.-M. Wang, *Angew. Chem., Int. Ed.*, 2023, **62**, e202304134.
- 32 G. Deng, H. Yun, M. S. Bootharaju, F. Sun, K. Lee, X. Liu, S. Yoo, Q. Tang, Y. J. Hwang and T. Hyeon, *J. Am. Chem. Soc.*, 2023, 145, 27407–27414.
- 33 G. Deng, J. Kim, M. S. Bootharaju, F. Sun, K. Lee, Q. Tang, Y. J. Hwang and T. Hyeon, *J. Am. Chem. Soc.*, 2023, 145, 3401–3407.
- 34 G. Deng, T. Ki, X. Liu, Y. Chen, K. Lee, S. Yoo, Q. Tang, M. S. Bootharaju and T. Hyeon, *Chem. Commun.*, 2024, 60, 1289–1292.
- 35 F. Hu, J.-J. Li, Z.-J. Guan, S.-F. Yuan and Q.-M. Wang, *Angew. Chem., Int. Ed.*, 2020, **59**, 5312–5315.
- 36 X. Wei, C. Xu, H. Li, X. Kang and M. Zhu, *Chem. Sci.*, 2022, 13, 5531–5538.
- 37 X. Ma, S. He, Q. Li, Q. Li, J. Chai, W. Ma, G. Li, H. Yu and M. Zhu, *Inorg. Chem.*, 2023, **62**, 15680–15687.
- 38 M. Bodiuzzaman, E. Khatun, K. S. Sugi, G. Paramasivam, W. A. Dar, S. Antharjanam and T. Pradeep, *J. Phys. Chem. C*, 2020, 124, 23426–23432.
- 39 B. J. Alamer, M. S. Bootharaju, S. M. Kozlov, Z. Cao, A. Shkurenko, S. Nematulloev, P. Maity, O. F. Mohammed, M. Eddaoudi, L. Cavallo, J.-M. Basset and O. M. Bakr, *Inorg. Chem.*, 2021, 60, 4306–4312.
- 40 L. G. AbdulHalim, M. S. Bootharaju, Q. Tang, S. Del Gobbo, R. G. AbdulHalim, M. Eddaoudi, D.-e. Jiang and O. M. Bakr, J. Am. Chem. Soc., 2015, 137, 11970–11975.
- 41 E. Khatun, A. Ghosh, P. Chakraborty, P. Singh, M. Bodiuzzaman, P. Ganesan, G. Nataranjan, J. Ghosh, S. K. Pal and T. Pradeep, *Nanoscale*, 2018, 10, 20033– 20042.
- 42 Q.-Q. Ma, X.-J. Zhai, J.-H. Huang, Y. Si, X.-Y. Dong, S.-Q. Zang and T. C. W. Mak, *Nanoscale*, 2024, **16**, 9361– 9366.
- 43 Y. Li, X.-M. Luo, P. Luo, Q.-X. Zang, Z.-Y. Wang and S.-Q. Zang, *ACS Nano*, 2023, 17, 5834–5841.

- 44 H. Li, X. Wei, X. Kang and M. Zhu, Nanoscale, 2024, 16, 1254-1259.
- 45 Z. Wang, H. Zhao, Y.-Z. Li, C. Zhang, R. K. Gupta, C.-H. Tung and D. Sun, *Nano Lett.*, 2024, 24, 458–465.
- 46 Z. Wang, F. Alkan, C. M. Aikens, M. Kurmoo, Z.-Y. Zhang, K.-P. Song, C.-H. Tung and D. Sun, Angew. Chem., Int. Ed., 2022, **61**, e202206742.
- 47 W.-J. Zhang, Z. Liu, K.-P. Song, C. M. Aikens, S.-S. Zhang, Z. Wang, C.-H. Tung and D. Sun, Angew. Chem., Int. Ed., 2021, 60, 4231-4237.
- 48 M. S. Bootharaju, H. Chang, G. Deng, S. Malola, W. Baek, H. Häkkinen, N. Zheng and T. Hyeon, J. Am. Chem. Soc., 2019, 141, 8422-8425.
- 49 M. Ou, H. Li, L.-H. Xie, S.-T. Yan, J.-R. Li, J.-H. Wang, C.-Y. Wei, Y.-W. Wu and X.-M. Zhang, J. Am. Chem. Soc., 2017, 139, 12346-12349.
- 50 M. S. Bootharaju, W. Baek, G. Deng, K. Singh, O. Voznyy, N. Zheng and T. Hyeon, Chem, 2022, 8, 2978-2989.
- 51 A. Gonzàlez-Rosell, S. Malola, R. Guha, N. R. Arevalos, M. F. Matus, M. E. Goulet, E. Haapaniemi, B. B. Katz, T. Vosch, J. Kondo, H. Häkkinen and S. M. Copp, J. Am. Chem. Soc., 2023, 145, 10721-10729.
- 52 L. Maretti, P. S. Billone, Y. Liu and J. C. Scaiano, J. Am. Chem. Soc., 2009, 131, 13972-13980.
- 53 Y.-X. Wang, J. Zhang, H.-F. Su, X. Cui, C.-Y. Wei, H. Li and X.-M. Zhang, ACS Nano, 2023, 17, 11607-11615.
- 54 C. P. Joshi, M. S. Bootharaju, M. J. Alhilaly and O. M. Bakr, J. Am. Chem. Soc., 2015, 137, 11578-11581.
- 55 F. Tian and R. Chen, J. Am. Chem. Soc., 2019, 141, 7107-7114.
- 56 M. S. Bootharaju, S. Lee, G. Deng, S. Malola, W. Baek, H. Häkkinen, N. Zheng and T. Hyeon, Angew. Chem., Int. Ed., 2021, 60, 9038-9044.
- 57 K.-G. Liu, X.-M. Gao, T. Liu, M.-L. Hu and D.-e. Jiang, J. Am. Chem. Soc., 2020, 142, 16905-16909.
- 58 H.-H. Wang, J. Wei, F. Bigdeli, F. Rouhani, H.-F. Su, L.-X. Wang, S. Kahlal, J.-F. Halet, J.-Y. Saillard, A. Morsali and K.-G. Liu, Nanoscale, 2023, 15, 8245-8254.
- 59 S.-F. Yuan, C.-Q. Xu, W.-D. Liu, J.-X. Zhang, J. Li and Q.-M. Wang, J. Am. Chem. Soc., 2021, 143, 12261-12267.
- 60 F. Hu, H.-W. Luyang, R.-L. He, Z.-J. Guan, S.-F. Yuan and Q.-M. Wang, J. Am. Chem. Soc., 2022, 144, 19365-19371.
- 61 G.-X. Duan, J. Han, B.-Z. Yang, Y.-P. Xie and X. Lu, Nanoscale, 2020, 12, 1617-1622.
- 62 S.-S. Zhang, F. Alkan, H.-F. Su, C. M. Aikens, C.-H. Tung and D. Sun, J. Am. Chem. Soc., 2019, 141, 4460-4467.
- 63 J. Wei, F. Bigdeli, L.-X. Wang, L.-L. Hou, A. Panjehpour, Y. Ma, K.-Z. Wang, A. Morsali and K.-G. Liu, Inorg. Chem., 2023, 62, 10185-10192.
- 64 M. Diecke, C. Schrenk and A. Schnepf, Angew. Chem., Int. Ed., 2020, **59**, 14418-14422.
- 65 C. Xu, Q. Yuan, X. Wei, H. Li, H. Shen, X. Kang and M. Zhu, Chem. Sci., 2022, 13, 1382-1389.
- 66 M.-X. Ma, X.-L. Ma, G.-M. Liang, X.-T. Shen, Q.-L. Ni, L.-C. Gui, X.-J. Wang, S.-Y. Huang and S.-M. Li, J. Am. Chem. Soc., 2021, 143, 13731-13737.

67 A. Jana, P. M. Unnikrishnan, A. K. Poonia, J. Roy, M. Jash, G. Paramasivam, J. Machacek, K. N. V. D. Adarsh, T. Base and T. Pradeep, Inorg. Chem., 2022, 61, 8593-8603.

- 68 L. Qin, F. Zhang, X. Ma, Y. Tang, G. Ma and Z. Tang, Dalton Trans., 2021, 50, 562-567.
- 69 J. Yang, S. Xie, H. Zhang, W. Xu, A. Dong and Y. Tang, Chem. Commun., 2022, 58, 6849-6852.
- 70 J. Yang and R. Jin, J. Phys. Chem. C, 2021, 125, 2619-2625.
- 71 X. Kang and M. Zhu, Chem. Soc. Rev., 2019, 48, 2422-
- 72 S. Wang, X. Meng, A. Das, T. Li, Y. Song, T. Cao, X. Zhu, M. Zhu and R. Jin, Angew. Chem., Int. Ed., 2014, 53, 2376-
- 73 X. Wang, B. Yin, L. Jiang, C. Yang, Y. Liu, G. Zou, S. Chen and M. Zhu, Science, 2023, 381, 784-790.
- 74 S. Chen, W. Du, C. Qin, D. Liu, L. Tang, Y. Liu, S. Wang and M. Zhu, Angew. Chem., Int. Ed., 2020, 59, 7542-7547.
- 75 M. S. Bootharaju, S. M. Kozlov, Z. Cao, A. Shkurenko, A. M. El-Zohry, O. F. Mohammed, M. Eddaoudi, O. M. Bakr, L. Cavallo and J.-M. Basset, Chem. Mater., 2018, 30, 2719-2725.
- 76 H. Shen, J. Xu, Z. Fu, X. Wei, X. Kang, W. Shi and M. Zhu, Angew. Chem., Int. Ed., 2024, 63, e202317995.
- 77 W. Ishii, Y. Okayasu, Y. Kobayashi, R. Tanaka, S. Katao, Y. Nishikawa, T. Kawai and T. Nakashima, J. Am. Chem. Soc., 2023, 145, 11236-11244.
- 78 H. Yang, J. Lei, B. Wu, Y. Wang, M. Zhou, A. Xia, L. Zheng and N. Zheng, Chem. Commun., 2013, 49, 300-302.
- 79 A. K. Das, R. Mekkat, S. Maity, A. S. Nair, S. Bhandary, R. Bhowal, A. Patra, B. Pathak, D. Chopra and S. Mandal, Inorg. Chem., 2021, 60, 19270-19277.
- 80 S. Kolay, S. Maity, S. Chakraborty, S. Ghosh and A. Patra, J. Phys. Chem. C, 2023, 127, 3769-3777.
- 81 S. Kolay, S. Chakraborty, S. Pramanik and A. Patra, J. Phys. Chem. C, 2024, 128, 7643-7651.
- 82 Z. Liu, M. Zhou, L. Luo, Y. Wang, E. Kahng and R. Jin, J. Am. Chem. Soc., 2023, 145, 19969-19981.
- 83 M. S. Bootharaju, C. P. Joshi, M. R. Parida, O. F. Mohammed and O. M. Bakr, Angew. Chem., Int. Ed., 2016, 55, 922-926.
- 84 K. S. Sugi, A. P. Sandra, Nonappa, D. Ghosh, J. S. Mohanty, M. Paulthangam Kannan, B. S. Sooraj, P. Srikrishnarka, J. Roy, W. A. Dar and T. Pradeep, Nanoscale, 2023, 15, 11927-11934.
- 85 M.-M. Zhang, X.-Y. Dong, Z.-Y. Wang, X.-M. Luo, J.-H. Huang, S.-Q. Zang and T. C. W. Mak, J. Am. Chem. Soc., 2021, 143, 6048-6053.
- 86 A. Jana, M. Jash, A. K. Poonia, G. Paramasivam, M. R. Islam, P. Chakraborty, S. Antharjanam, J. Machacek, S. Ghosh, K. N. V. D. Adarsh, T. Base and T. Pradeep, ACS Nano, 2021, 15, 15781-15793.
- 87 A. Jana, W. A. Dar, S. K. Jana, A. K. Poonia, V. Yadav, J. Roy, S. Chandra, K. N. V. D. Adarsh, R. H. A. Ras and T. Pradeep, Chem. Mater., 2023, 35, 7020-7031.
- 88 T. Chen, S. Yang, Y. Song, J. Chai, Q. Li, X. Ma, G. Li, H. Yu and M. Zhu, Chem. Commun., 2020, 56, 7605-7608.

- 89 S. Yoo, S. Yoo, G. Deng, F. Sun, K. Lee, H. Jang, C. W. Lee, X. Liu, J. Jang, Q. Tang, Y. J. Hwang, T. Hyeon and M. S. Bootharaju, Adv. Mater., 2024, 36, 2313032.
- 90 A. Jana, P. Chakraborty, W. A. Dar, S. Chandra, E. Khatun, M. P. Kannan, R. H. A. Ras and T. Pradeep, Chem. Commun., 2020, 56, 12550-12553.
- 91 Y.-M. Tseng, J.-H. Liao, T.-H. Chiu, H. Liang, S. Kahlal, J.-Y. Saillard and C. W. Liu, Inorg. Chem., 2023, 62, 3866-3874.
- 92 A. Baksi, M. S. Bootharaju, P. K. Chhotaray, P. Chakraborty, B. Mondal, S. Bhat, R. Gardas and T. Pradeep, J. Phys. Chem. C, 2017, 121, 26483-26492.
- 93 M. S. Bootharaju, S. M. Kozlov, Z. Cao, M. Harb, N. Maity, A. Shkurenko, M. R. Parida, M. N. Hedhili, M. Eddaoudi, O. F. Mohammed, O. M. Bakr, L. Cavallo and J.-M. Basset, J. Am. Chem. Soc., 2017, 139, 1053-1056.
- 94 M. S. Bootharaju, S. Lee, G. Deng, H. Chang, W. Baek and T. Hyeon, J. Chem. Phys., 2021, 155, 014307.
- 95 W.-J. Yen, J.-H. Liao, T.-H. Chiu, Y.-S. Wen and C. W. Liu, Inorg. Chem., 2024, 63, 5320-5324.
- 96 Y.-J. Zhong, J.-H. Liao, T.-H. Chiu, S. Kahlal, C.-J. Lin, J.-Y. Saillard and C. W. Liu, Angew. Chem., Int. Ed., 2021, 60, 12712-12716.
- 97 X. Kang, X. Wei, P. Xiang, X. Tian, Z. Zuo, F. Song, S. Wang and M. Zhu, Chem. Sci., 2020, 11, 4808-4816.
- 98 G. Deng, S. Malola, P. Yuan, X. Liu, B. K. Teo, H. Häkkinen and N. Zheng, Angew. Chem., Int. Ed., 2021, 60, 12897-12903.
- 99 G. Deng, B. K. Teo and N. Zheng, J. Am. Chem. Soc., 2021, 143, 10214-10220.
- 100 S. Wang, Y. Tan, T. Li, Q. Zhou, P. Li, S. Yang, H. Yu and M. Zhu, Inorg. Chem., 2022, 61, 18450-18457.
- 101 L. Qin, F. Sun, X. Ma, G. Ma, Y. Tang, L. Wang, Q. Tang, R. Jin and Z. Tang, Angew. Chem., Int. Ed., 2021, 60, 26136-26141.
- 102 L. Chen, F. Sun, Q. Shen, L. Qin, Y. Liu, L. Qiao, Q. Tang, L. Wang and Z. Tang, Nano Res., 2022, 15, 8908-8913.
- 103 L.-J. Li, Y.-T. Luo, Y.-Q. Tian, P. Wang, X.-Y. Yi, J. Yan, Y. Pei and C. Liu, *Inorg. Chem.*, 2023, **62**, 14377–14384.
- 104 Y. Xiong, H. Chen, Y. Hu, S. Yang, X. Xue, L. He, X. Liu, J. Ma and Z. Jin, *Nano Lett.*, 2021, 21, 8693–8700.
- 105 Y. Feng, F. Fu, L. Zeng, M. Zhao, X. Xin, J. Liang, M. Zhou, X. Fang, H. Lv and G.-Y. Yang, Angew. Chem., Int. Ed., 2024, **63**, e202317341.
- 106 Y. Wang, X.-H. Liu, Q. Wang, M. Quick, S. A. Kovalenko, Q.-Y. Chen, N. Koch and N. Pinna, Angew. Chem., Int. Ed., 2020, 59, 7748-7754.

- 107 M. S. Bootharaju, C. W. Lee, G. Deng, H. Kim, K. Lee, S. Lee, H. Chang, S. Lee, Y.-E. Sung, J. S. Yoo, N. Zheng and T. Hyeon, Adv. Mater., 2023, 35, 2207765.
- 108 Y. Liu, D. Long, A. Springer, R. Wang, N. Koch, M. Schwalbe, N. Pinna and Y. Wang, Sol. RRL, 2023, 7, 2201057.
- 109 H. Shen, Q. Zhu, J. Xu, K. Ni, X. Wei, Y. Du, S. Gao, X. Kang and M. Zhu, Nanoscale, 2023, 15, 14941–14948.
- 110 L. Liu, S.-J. Zheng, H. Chen, J. Cai and S.-Q. Zang, Angew. Chem., Int. Ed., 2024, 63, e202316910.
- 111 L. Qin, F. Sun, Z. Gong, G. Ma, Y. Chen, Q. Tang, L. Qiao, R. Wang, Z.-Q. Liu and Z. Tang, ACS Nano, 2023, 17, 12747-12758.
- 112 K. Yonesato, S. Yamazoe, D. Yokogawa, K. Yamaguchi and K. Suzuki, Angew. Chem., Int. Ed., 2021, 60, 16994-16998.
- 113 J. Zhuge, F. Rouhani, F. Bigdeli, X.-M. Gao, H. Kaviani, H.-J. Li, W. Wang, M.-L. Hu, K.-G. Liu and A. Morsali, Dalton Trans., 2021, 50, 2606-2615.
- 114 Q.-Y. Wang, J. Wang, S. Wang, Z.-Y. Wang, M. Cao, C.-L. He, J.-Q. Yang, S.-Q. Zang and T. C. W. Mak, J. Am. Chem. Soc., 2020, 142, 12010-12014.
- 115 C. Wang, Y.-J. Wang, C.-L. He, Q.-Y. Wang and S.-Q. Zang, JACS Au, 2021, 1, 2202-2207.
- 116 H. Li, W. Duan, X. Cui, S. Liu, Y. Bai, Y.-X. Wang, S. Feng, Y. Wang, M. Qu and X.-M. Zhang, ACS Appl. Nano Mater., 2022, 5, 14251-14255.
- 117 J.-Y. Wang, Y.-K. Li, X. Jing, P. Luo, X.-Y. Dong and S.-Q. Zang, ACS Mater. Lett., 2022, 4, 960–966.
- 118 Z.-J. Guan, R.-L. He, S.-F. Yuan, J.-J. Li, F. Hu, C.-Y. Liu and Q.-M. Wang, Angew. Chem., Int. Ed., 2022, 61, e202116965.
- 119 J. Yuan, X. Huang, W. Zhang, M. Zhou, G. Li, F. Tian and R. Chen, Inorg. Chem., 2023, 62, 17668-17677.
- 120 K. Yonesato, D. Yanai, S. Yamazoe, D. Yokogawa, T. Kikuchi, K. Yamaguchi and K. Suzuki, Nat. Chem., 2023, 15, 940-947.
- 121 K. O. Sulaiman, A. Bueckert, A. Abdellah, S. K. Veeranmaril, D. C. Higgins and R. W. J. Scott, J. Phys. Chem. C, 2022, **126**, 16117–16126.
- 122 K. O. Sulaiman, V. Sudheeshkumar and R. W. J. Scott, RSC *Adv.*, 2019, **9**, 28019–28027.
- 123 L. Li, Y. Lv, H. Sheng, Y. Du, H. Li, Y. Yun, Z. Zhang, H. Yu and M. Zhu, Nat. Commun., 2023, 14, 6989.
- 124 T. U. B. Rao, B. Nataraju and T. Pradeep, J. Am. Chem. Soc., 2010, 132, 16304-16307.
- 125 T. Udayabhaskararao, M. S. Bootharaju and T. Pradeep, Nanoscale, 2013, 5, 9404-9411.