


Cite this: *J. Mater. Chem. A*, 2017, 5, 14530Received 29th April 2017
Accepted 22nd June 2017

DOI: 10.1039/c7ta03703a

rsc.li/materials-a

Hierarchical NiCo₂O₄ nanosheets on carbon nanofiber films for high energy density and long-life Li–O₂ batteries†

Guoxue Liu, Lei Zhang,* Suqing Wang, Liang-Xin Ding and Haihui Wang *

Designing oxygen cathodes with both high energy density and excellent cycling stability is a great challenge in the development of lithium–oxygen (Li–O₂) batteries for energy storage systems. Herein, we design a novel structure of hierarchical NiCo₂O₄ nanosheets on porous carbon nanofiber films (denoted as NiCo₂O₄@CNFs) as an oxygen cathode for lithium–oxygen batteries. The NiCo₂O₄@CNFs cathode delivers a high specific discharge capacity of 4179 mA h g^{−1}, a high energy density of 2110 W h kg^{−1} and superior cycling stability over 350 cycles. The excellent electrochemical performance of the NiCo₂O₄@CNFs cathode can be attributed to the rational design and engineering of catalysts and porous conductive electrodes. These results indicate that the NiCo₂O₄@CNFs electrode is a promising candidate for high energy density and long-life Li–O₂ batteries. Additionally, the rational design of the hierarchical catalyst constructed low-dimensional nanostructure and the lightweight porous carbon nanofiber electrode can be also used for other metal–oxygen batteries, such as zinc–oxygen (Zn–O₂) batteries, aluminum–oxygen (Al–O₂) batteries, and sodium–oxygen (Na–O₂) batteries.

Rechargeable lithium–oxygen (Li–O₂) batteries are unmatched candidates amongst future electrochemical energy storage devices for electric vehicles due to their high theoretical energy density of 3505 W h kg^{−1}, which is about 10 times higher than that of conventional Li-ion batteries (360 W h kg^{−1}).^{1–6} In a typical prototype, a Li–O₂ battery is composed of an Li⁺ conducting electrolyte, a separator, a lithium anode, and a porous O₂ breathing cathode, and its cycling mechanism is based on the oxygen reduction reaction (ORR) and the oxygen evolution reaction (OER) between Li⁺ and oxygen on the cathode: 2Li⁺ + O₂ + 2e[−] ↔ Li₂O₂.^{7,8} As a result, the key of the development of high-performance Li–O₂ batteries depends on the cathode, which needs to meet such requirements: highly active bifunctional ORR/OER catalysts, and porous conductive electrodes for

oxygen diffusion and electron transport, as well as offering enough space to meet the needs of Li₂O₂ deposition.^{9–12}

Despite their superior theoretical storage capacity, constructing cycling stable Li–O₂ battery cathodes is still facing materials challenges.^{11,13} This is traced to the sluggish kinetics of the ORR/OER giving large discharge/charge overpotentials and the random deposition of Li₂O₂ blocking the permeation path for oxygen and the electrolyte, leading to low round-trip efficiency, poor rate capability, and short cycling performance.^{9,14} From the catalyst point of view, although the most efficient catalysts for the sluggish bifunctional ORR/OER are still noble metals,^{15–19} low cost earth-abundant transition metal oxides, such as MnO₂,^{20–23} NiCo₂O₄,^{24,25} Co₃O₄,^{26–28} MnCo₂O₄,^{29,30} and CoMn₂O₄,^{31,32} are preferred for practical application. Among these, NiCo₂O₄-based materials have received considerable attention due to their rich redox reaction sites from both redox couples of Co³⁺/Co²⁺ and Ni³⁺/Ni²⁺, good electrical conductivity, and affordable cost.^{33–35} Hence, low-dimensional nanostructured NiCo₂O₄-based catalysts, such as 0D nanoparticles, 1D nanorods/nanowires/nanotubes and 2D nanosheets, have been reported as bifunctional ORR/OER catalysts for rechargeable Li–O₂ batteries.^{36–41} Low dimensional NiCo₂O₄ nanostructures possess highly exposed active sites and short diffusion length for the ORR/OER, which can lead to a high discharge capacity and low overpotentials.³⁶ However, these low-dimensional nanostructures cannot ensure enough space for solid Li₂O₂ deposition, which could cover catalytically active sites and result in poor cycling performance.²⁴ Also, the mass transport channels of Li⁺ and oxygen are easily clogged due to relatively small pores of low-dimensional nanostructures, leading to cell failure.⁴² Therefore, the fabrication of NiCo₂O₄-based catalysts with efficiently sufficient space and an accessible porous structure is highly desirable for Li–O₂ batteries.

As to the porous conductive electrode, a “binder-free” cathode has recently been introduced as an ideal electrode to avoid the side reaction between the binder and electrolyte.⁴³ Among the substrates, Ni foam and carbon textile are most commonly used because of their good breathability, high

School of Chemistry and Chemical Engineering, South China University of Technology, Guangzhou 510640, China. E-mail: celeizhang@scut.edu.cn; hhwang@scut.edu.cn

† Electronic supplementary information (ESI) available. See DOI: 10.1039/c7ta03703a

conductivity and good stability. For example, Liu *et al.* synthesized a cathode of $\delta\text{-MnO}_2$ assembled by ultrathin nanosheets on Ni foam coated with graphene, which showed excellent stability to enhance the cycling stability of Li–O₂ batteries.²³ Zhang *et al.* described that TiO₂ nanowires grown on carbon clothes exhibited good electrochemical performance attributed to their high mechanical and chemical stability.⁴⁴ However, the high mass densities of Ni foam and carbon textile will lead to sacrificing capacity and energy density.³⁹ Thus, it is highly demanded to design and develop a lightweight cathode with high catalytic ORR/OER activities for realizing the high energy density of Li–O₂ batteries.

To achieve a Li–O₂ battery with high cycling stability and high energy density, it is highly necessary to combine a rationally structured catalyst and a lightweight porous conductive electrode. Herein, we propose a novel structure of hierarchical NiCo₂O₄ nanosheets on porous carbon nanofiber films (denoted as NiCo₂O₄@CNFs), for high energy density and long-life Li–O₂ batteries. The nanostructured cathode integrates several desirable design rationales for high-performance Li–O₂ batteries based on low-dimensional ultrathin nanosheets, lightweight conductive carbon networks, and a binder-free cathode. With the help of this rational design, the NiCo₂O₄@CNFs cathode exhibits excellent electrochemical performance including a high specific discharge capacity of 4179 mA h g⁻¹, an excellent energy density of 2110 W h kg⁻¹, and especially superior cycling stability over 350 cycles.

The fabrication procedure of porous NiCo₂O₄ nanosheets on carbon nanofiber films and the structure of the Li–O₂ battery are described in Fig. 1. Typically, porous carbon nanofiber films were prepared *via* a facile electrospinning technique followed by carbonization. In step I, Ni–Co precursor nanosheets were grown on carbon nanofiber films by a hydrothermal method to form core–shell structures. In step II, the Ni–Co precursor nanosheets were transformed into porous NiCo₂O₄ nanosheets through annealing and the NiCo₂O₄@CNFs cathode was finally

obtained. The morphology and microstructure of the cathode were investigated by field emission scanning electron microscopy (FESEM). As shown in Fig. 2, the randomly arranged carbon nanofibers form conductive networks with numerous irregular micrometer-scale pores between the carbon nanofibers, which could enhance electron transfer among cathodes and offer sufficient channels for cathode breathing. From the high magnification SEM, it can be seen that the as-prepared carbon nanofibers have an average diameter of about 180 nm with a smooth surface (Fig. 2b). The X-ray diffraction (XRD) pattern of CNFs shows no obvious diffraction peaks, indicating that the carbon nanofibers consist of typical amorphous carbon (Fig. S1, ESI†). Afterwards, Ni–Co precursors were grown on carbon nanofibers through a simple hydrothermal method. As shown in Fig. 2c and d, the low-magnification FESEM image of NiCo₂O₄@CNFs (Fig. 2c) displays that the Ni–Co precursor nanosheets uniformly coat the carbon nanofibers and the pores between the individual carbon nanofibers are not been covered. The magnified FESEM image (Fig. 2d) shows that the Ni–Co precursor nanosheets are around 10 nm thick. After annealing at 350 °C in a nitrogen atmosphere, the Ni–Co precursor nanosheets were converted to NiCo₂O₄ nanosheets. From the FESEM images (Fig. 2e), it can be clearly observed that the NiCo₂O₄ shell is still constructed by nanosheet-like subunits on the carbon nanofibers because of the robust support of the carbon nanofibers. Interestingly, from the FESEM image of a single 1D NiCo₂O₄@CNFs (Fig. 2f), the NiCo₂O₄ nanosheets are clearly observed with a rough and porous surface.

The crystallographic phase of the NiCo₂O₄ nanosheets was characterized by X-ray powder diffraction (XRD), as shown in Fig. S2 (ESI†). All of the diffraction peaks in the XRD pattern can be unambiguously assigned to the spinel NiCo₂O₄ (JCPDS no. 20-0781).^{24,34} No other diffraction peaks from possible impurities are observed, suggesting that the Ni–Co precursor is converted into highly pure spinel NiCo₂O₄. By thermogravimetric analysis (TGA, Fig. S3, ESI†), the mass content of NiCo₂O₄ in NiCo₂O₄@CNFs is approximately 40%.^{36,39} The chemical

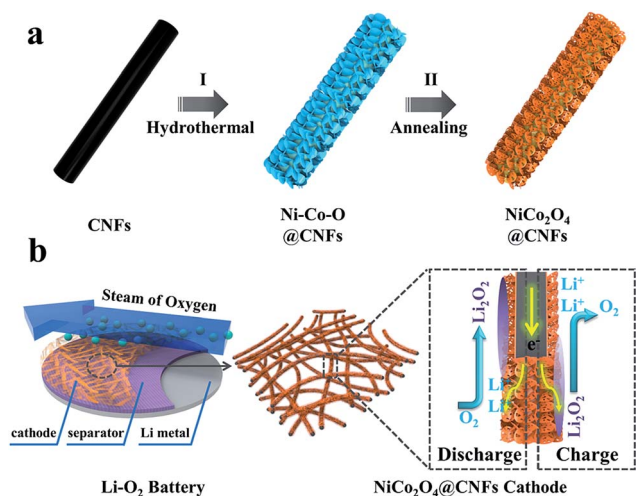


Fig. 1 (a) Schematic illustration of the fabrication of the NiCo₂O₄@CNFs cathode; (b) structure of the rechargeable Li–O₂ battery containing this cathode.

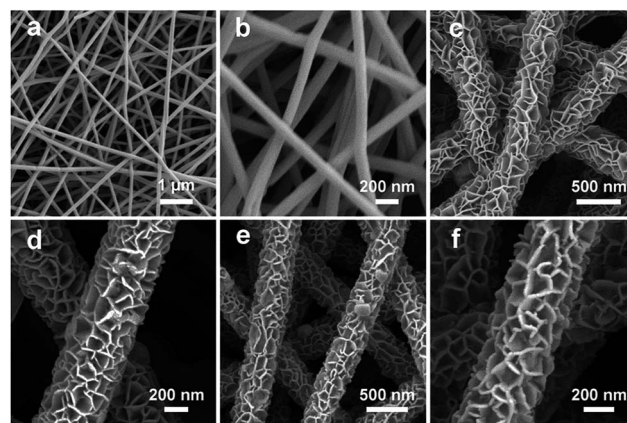


Fig. 2 FESEM images of (a, b) carbon nanofiber films at different magnifications; (c, d) Ni–Co–O precursor nanosheets grown on carbon nanofibers at different magnifications; (e, f) NiCo₂O₄ nanosheets grown on carbon nanofibers at different magnifications.

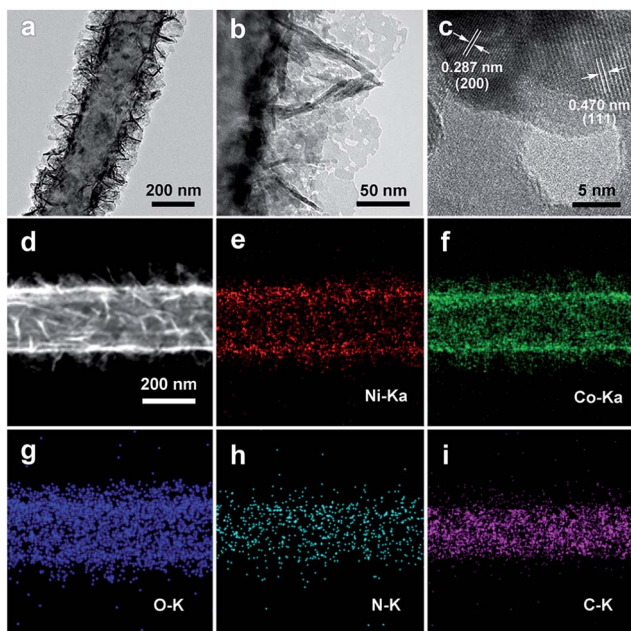


Fig. 3 (a, b) Typical TEM images of NiCo_2O_4 nanosheets on carbon nanofiber films at different magnifications; (c) an HRTEM image of NiCo_2O_4 nanosheets; (d–i) HAADF-STEM and corresponding EDX elemental mappings (Ni, Co, O, N, C) of NiCo_2O_4 @CNFs.

compositions of NiCo_2O_4 @CNFs were further analyzed by energy dispersive X-ray (EDX) spectroscopy (Fig. S4, ESI[†]). The total mass of Ni, Co, and O is 43.8%, which is slightly higher than the TGA result, which might be caused by the oxygen atoms of CNFs. In virtue of 1D nanofibers and the ultrathin sheet-like subunits, these NiCo_2O_4 @CNFs have a relatively large Brunauer–Emmett–Teller (BET) surface area of about $178 \text{ m}^2 \text{ g}^{-1}$ (N_2 adsorption desorption isotherm is given in Fig. S5, ESI[†]).

To further understand the porous nanosheet structure, transmission electron microscopy (TEM) was performed (Fig. 3). In Fig. 3a and b, it can be clearly observed that NiCo_2O_4 nanosheets are uniformly grown on the CNFs (Fig. 3a), and the surface of NiCo_2O_4 contains abundant mesopores (Fig. 3b). Furthermore, the high-resolution TEM (HRTEM, Fig. 3c) images of NiCo_2O_4 nanosheets exhibits lattice fringes with inter-planar spacings of 0.287 nm and 0.470 nm, corresponding to the (200) and (111) planes of NiCo_2O_4 . Energy-dispersive X-ray (EDX) mapping was used to further analyze the distribution of the catalyst on CNFs, as shown in Fig. 3d–i. The Ni, Co, and O elements are uniformly distributed around a carbon nanofiber, and the C element is distributed into the core of the composite. In addition, energy dispersive X-ray (EDX, Fig. S6, ESI[†]) spectral line-scan results show that the C element is mainly detected

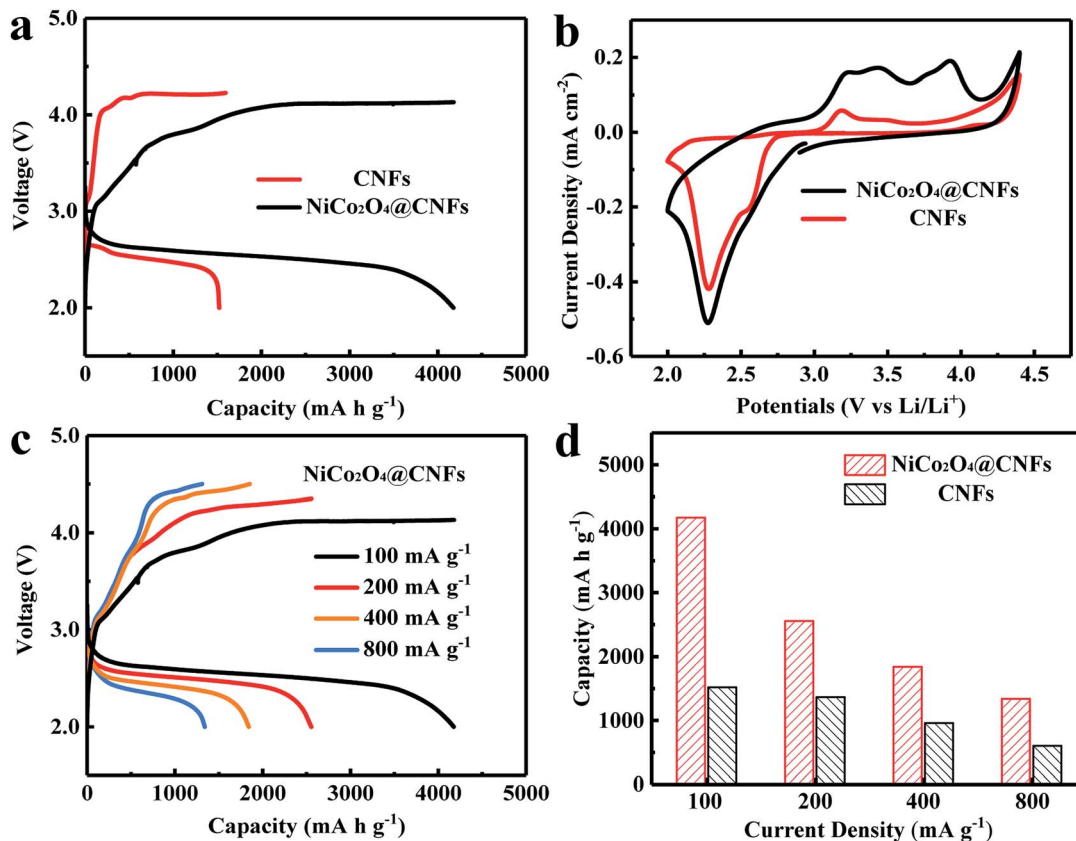


Fig. 4 (a) Initial discharge/charge curves of Li–O₂ batteries with NiCo_2O_4 @CNFs and CNFs cathodes at a current density of 100 mA g^{-1} ; (b) CV curves of NiCo_2O_4 @CNFs and CNFs cathodes between 2.0 and 4.5 V at 0.1 mV s^{-1} ; (c) discharge/charge curves of Li–O₂ batteries with the NiCo_2O_4 @CNFs cathode at various current densities; (d) the rate capability of Li–O₂ batteries with NiCo_2O_4 @CNFs and CNFs cathodes at various current densities.

inside the NiCo₂O₄@CNFs, and the Ni, Co, and O elements are mainly detected on the shell, which could suggest that the CNFs are covered by a uniform coating layer of NiCo₂O₄.⁴⁵

We subsequently investigated the electrochemical properties of the NiCo₂O₄@CNFs as a cathode for Li–O₂ batteries. Fig. 4a shows the discharge–charge voltage profiles of NiCo₂O₄@CNFs and CNFs cathodes at a current density of 100 mA g^{−1}. The NiCo₂O₄@CNFs cathode shows a high initial discharge capacity of 4179 mA h g^{−1}, which is higher than that of the CNFs cathode (1591 mA h g^{−1}). Furthermore, the NiCo₂O₄@CNFs cathode presents lower overpotentials than CNFs, which suggests that the NiCo₂O₄ catalyst has higher electrocatalytic activities toward the ORR/OER.^{46,47} The NiCo₂O₄@CNFs cathode with different loading masses at the same mass current density was investigated as shown in Fig. S7.† The specific capacities of the NiCo₂O₄@CNFs cathode with loading masses of 0.134, 0.268, and 0.536 mg cm^{−2} were 4179, 2231, and 1992 mA h g^{−1}, respectively. The morphological changes of the NiCo₂O₄@CNFs cathode after the 1st full discharge and 1st charge are examined (Fig. S8, ESI†). As shown Fig. S8b, ESI†, the NiCo₂O₄ shell is constructed by nanosheet-like subunits on the carbon nanofibers. The part of 3D spaces with NiCo₂O₄ nanosheets was filled with the discharge product (Li₂O₂) during the discharge process from 0 mA h g^{−1} to 1000 mA h g^{−1} (Fig. S8c, ESI†). Then, the 3D structure was completely covered by the discharge product after being discharged to 2.0 V (Fig. S8d, ESI†). After the charge

process, the 3D structure consisting of porous NiCo₂O₄ nanosheets was still observed obviously (Fig. S8e, ESI†) suggesting that the NiCo₂O₄@CNFs can remain stable enough during the discharge and charge processes and the discharge product inside the 3D structure was decomposed completely after charging, which suggest that the NiCo₂O₄@CNFs cathode show good mass transfer channels during the discharge and charge processes enhancing its capacity and recovery.⁴⁷ In order to further investigate the electrocatalytic activities of the NiCo₂O₄@CNFs, cyclic voltammetry (CV) was performed at a scan rate of 0.1 mV s^{−1} in the voltage window of 2.0–4.5 V, as shown in Fig. 4b. The CV behaviors are similar to those previously reported, suggesting that the oxygen electrodes have the same reaction pathway.^{23,46,48} Specifically, one reduction peak can be clearly observed, which can be ascribed to the oxygen reduction. In addition, the NiCo₂O₄@CNFs cathode exhibits a lower onset ORR potential and larger ORR peak current density than CNFs. Consistent with the result of the discharge–charge profile, the CV curve of the NiCo₂O₄@CNFs cathode shows three oxidation peaks at around 3.24, 3.44, and 3.93 V, respectively, which is attributed to the stepwise process of decomposition of Li₂O₂.^{14,48,49} Moreover, the NiCo₂O₄@CNFs cathode shows a higher OER peak current density than CNFs. The rate capability of the NiCo₂O₄@CNFs and CNFs cathodes was investigated through galvanostatic discharge–charge at various current densities from 100 to 800 mA g^{−1}, as shown in Fig. 4c

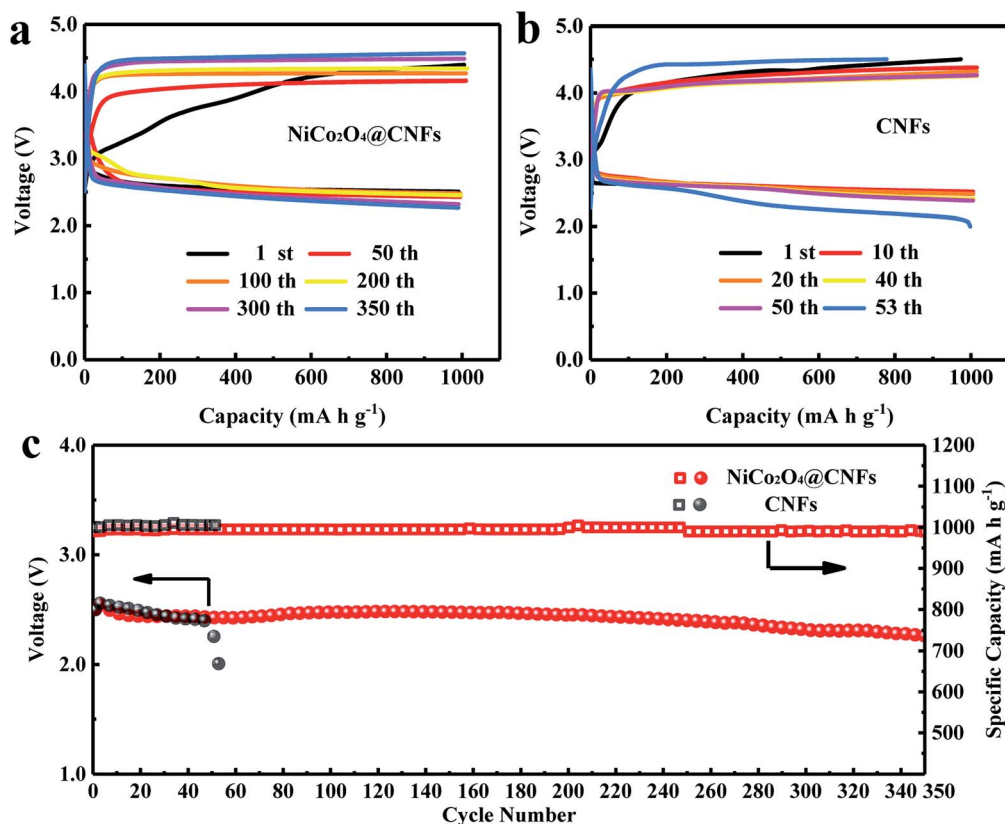


Fig. 5 (a and b) The cycling performance of NiCo₂O₄@CNFs and CNFs cathodes at a current density of 200 mA g^{−1} and a specific capacity limit of 1000 mA h g^{−1}; (c) the terminal discharge voltage as a function of cycle number for Li–O₂ batteries with NiCo₂O₄@CNFs and CNFs cathodes.

high energy density and excellent cycling stability. From these results, the rational design of the hierarchical catalyst constructed low-dimensional nanostructure and the lightweight porous carbon nanofiber electrode toward oxygen cathodes can be also used for other metal–oxygen batteries, such as Zn–O₂ batteries, Al–O₂ batteries, and Na–O₂ batteries.

Acknowledgements

This work was financially supported by the National Key Research and Development Program of China (2016YFA0202600), the “Thousand Talents Program”, National Natural Science Foundation of China (21606088, 51621001, 21576100), and the Pearl River S&T Nova Program of Guangzhou (201610010062).

Notes and references

- P. G. Bruce, S. A. Freunberger, L. J. Hardwick and J. M. Tarascon, *Nat. Mater.*, 2012, **11**, 19.
- N. Feng, P. He and H. Zhou, *Adv. Energy Mater.*, 2016, **6**, 201502303.
- Y. Y. Shao, F. Ding, J. Xiao, J. Zhang, W. Xu, S. Park, J. G. Zhang, Y. Wang and J. Liu, *Adv. Funct. Mater.*, 2013, **23**, 987.
- Y. Li, X. Wang, S. Dong, X. Chen and G. Cui, *Adv. Energy Mater.*, 2016, **6**, 201600751.
- G. Girishkumar, B. McCloskey, A. C. Luntz, S. Swanson and W. Wilcke, *J. Phys. Chem. Lett.*, 2010, **1**, 2193.
- J. Lu, L. Li, J. B. Park, Y. K. Sun, F. Wu and K. Amine, *Chem. Rev.*, 2014, **114**, 5611.
- W. B. Luo, X. W. Gao, D. Q. Shi, S. L. Chou, J. Z. Wang and H. K. Liu, *Small*, 2016, **12**, 3031.
- K. Liao, X. Wang, Y. Sun, D. Tang, M. Han, P. He, X. Jiang, T. Zhang and H. Zhou, *Energy Environ. Sci.*, 2015, **8**, 1992.
- L. Grande, E. Paillard, J. Hassoun, J. B. Park, Y. J. Lee, Y. K. Sun, S. Passerini and B. Scrosati, *Adv. Mater.*, 2015, **27**, 784.
- J. Wang, Y. Li and X. Sun, *Nano Energy*, 2013, **2**, 443.
- Z. W. Chang, J. J. Xu, Q. C. Liu, L. Li and X. B. Zhang, *Adv. Energy Mater.*, 2015, **5**, 201500633.
- A. Kraytsberg and Y. Ein-Eli, *J. Power Sources*, 2011, **196**, 886.
- J. S. Lee, S. T. Kim, R. Cao, N. S. Choi, M. Liu, K. T. Lee and J. Cho, *Adv. Energy Mater.*, 2011, **1**, 34.
- Y. C. Lu, B. M. Gallant, D. G. Kwabi, J. R. Harding, R. R. Mitchell, M. S. Whittingham and S. H. Yang, *Energy Environ. Sci.*, 2013, **6**, 750.
- Z. Peng, S. A. Freunberger, Y. Chen and P. G. Bruce, *Science*, 2012, **337**, 563.
- D. Oh, J. F. Qi, Y. C. Lu, Y. Zhang, S. H. Yang and A. M. Belcher, *Nat. Commun.*, 2013, **4**, 2756.
- J. J. Xu, Z. L. Wang, D. Xu, L. L. Zhang and X. B. Zhang, *Nat. Commun.*, 2013, **4**, 2438.
- Y. C. Lu, Z. C. Xu, H. A. Gasteiger, S. Chen, H. S. Kimberly and S. H. Yang, *J. Am. Chem. Soc.*, 2010, **132**, 12170.
- F. J. Li, D. M. Tang, Z. L. Jian, D. Q. Liu, D. Golberg, A. Yamada and H. S. Zhou, *Adv. Mater.*, 2014, **26**, 4659.
- A. Débart, A. J. Paterson, J. Bao and P. G. Bruce, *Angew. Chem., Int. Ed.*, 2008, **120**, 4597.
- Y. Cao, Z. Wei, J. He, J. Zang, Q. Zhang, M. Zheng and Q. Dong, *Energy Environ. Sci.*, 2012, **5**, 9765.
- X. Hu, F. Cheng, X. Han, T. Zhang and J. Chen, *Small*, 2015, **11**, 809.
- S. Liu, Y. Zhu, J. Xie, Y. Huo, H. Y. Yang, T. Zhu, G. Cao, X. Zhao and S. Zhang, *Adv. Energy Mater.*, 2014, **4**, 201301960.
- L. L. Liu, J. Wang, Y. Y. Hou, J. Chen, H. K. Liu, J. Z. Wang and Y. P. Wu, *Small*, 2016, **12**, 602.
- B. Sun, X. Huang, S. Chen, Y. Zhao, J. Zhang, P. Munroe and G. Wang, *J. Mater. Chem. A*, 2014, **2**, 12053.
- Y. Liang, Y. Li, H. Wang, J. Zhou, J. Wang, T. Regier and H. Dai, *Nat. Mater.*, 2011, **10**, 780.
- W. H. Ryu, T. H. Yoon, S. H. Song, S. Jeon, Y. J. Park and I. D. Kim, *Nano Lett.*, 2013, **13**, 4190.
- T. Y. Ma, S. Dai, M. Jaroniec and S. Z. Qiao, *J. Am. Chem. Soc.*, 2014, **136**, 13925.
- D. Oh, J. F. Qi, B. H. Han, G. R. Zhang, T. J. Carney, J. Ohmura, Y. Zhang, S. H. Yang and A. M. Belcher, *Nano Lett.*, 2014, **14**, 4837.
- H. Wang, Y. Yang, Y. Liang, G. Zheng, Y. Li, Y. Cui and H. Dai, *Energy Environ. Sci.*, 2012, **5**, 7931.
- S. Peng, L. Li, Y. Hu, M. Srinivasan, F. Cheng, J. Chen and S. Ramakrishna, *ACS Nano*, 2015, **9**, 1945.
- C. Li, X. P. Han, F. Y. Cheng, Y. X. Hu, C. C. Chen and J. Chen, *Nat. Commun.*, 2015, **6**, 7345.
- S. X. Liu, L. F. Hu, X. J. Xu, A. A. Al-Ghamdi and X. S. Fang, *Small*, 2015, **11**, 4267.
- Z. Gao, W. Yang, J. Wang, N. Song and X. Li, *Nano Energy*, 2015, **13**, 306.
- G. Zhang and X. W. Lou, *Adv. Mater.*, 2013, **25**, 976.
- H. Gong, H. R. Xue, T. Wang, H. Guo, X. L. Fan, L. Song, W. Xia and J. P. He, *ACS Appl. Mater. Interfaces*, 2016, **8**, 18060.
- H. Xue, X. Mu, J. Tang, X. Fan, H. Gong, T. Wang, J. He and Y. Yamauchi, *J. Mater. Chem. A*, 2016, **4**, 9106.
- Y. Luo, F. L. Lu, C. Jin, Y. R. Wang, R. Z. Yang and C. H. Yang, *J. Power Sources*, 2016, **319**, 19.
- H. R. Xue, S. C. Wu, J. Tang, H. Gong, P. He, J. P. He and H. S. Zhou, *ACS Appl. Mater. Interfaces*, 2016, **8**, 8427.
- L. Y. Li, L. F. Shen, P. Nie, G. Pang, J. Wang, H. S. Li, S. Y. Dong and X. G. Zhang, *J. Mater. Chem. A*, 2015, **3**, 24309.
- A. Riaz, K. N. Jung, W. Chang, K. H. Shin and J. W. Lee, *ACS Appl. Mater. Interfaces*, 2014, **6**, 17815.
- J. Xie, X. Yao, Q. Cheng, I. P. Madden, P. Dornath, C. C. Chang, W. Fan and D. Wang, *Angew. Chem., Int. Ed.*, 2015, **54**, 4299.
- D. Xiao, S. Dong, J. Guan, L. Gu, S. Li, N. Zhao, C. Shang, Z. Yang, H. Zheng, C. Chen, R. Xiao, Y. S. Hu, H. Li, G. Cui and L. Chen, *Adv. Energy Mater.*, 2015, **5**, 1400664.
- Q. C. Liu, J. J. Xu, D. Xu and X. B. Zhang, *Nat. Commun.*, 2015, **6**, 7892.
- Z. L. Jian, P. Liu, F. J. Li, P. He, X. W. Guo, M. W. Chen and H. S. Zhou, *Angew. Chem., Int. Ed.*, 2014, **53**, 442.

- 46 G. Liu, H. Chen, L. Xia, S. Wang, L. X. Ding, D. Li, K. Xiao, S. Dai and H. Wang, *ACS Appl. Mater. Interfaces*, 2015, **7**, 22478.
- 47 J. J. Xu, D. Xu, Z. L. Wang, H. G. Wang, L. L. Zhang and X. B. Zhang, *Angew. Chem., Int. Ed.*, 2013, **52**, 3887.
- 48 F. F. Tu, J. Xie, S. C. Zhang, G. S. Cao, T. J. Zhu and X. B. Zhao, *J. Mater. Chem. A*, 2015, **3**, 5714.
- 49 D. Zhai, H. H. Wang, J. Yang, K. C. Lau, K. Li, K. Amine and L. A. Curtiss, *J. Am. Chem. Soc.*, 2013, **135**, 15364.
- 50 R. R. Mitchell, B. M. Gallant, C. V. Thompson and S. H. Yang, *Energy Environ. Sci.*, 2011, **4**, 2952.