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# High-loading metal atoms on graphdiyne for efficient nitrogen fixation to ammonia†

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Atomic catalysts have shown great potential in the field of catalysis because of their maximum atomic utilization, and high reaction selectivity and activity. Graphdiyne-based atomic catalysts have emerged as a new class of promising materials in catalysis. Here, a new graphdiyne-based manganese atomic electrocatalyst (Mn<sub>SA</sub>/GDY) was designed and successfully synthesized *via* a facile *in situ* reduction strategy for electrochemical nitrogen fixation to ammonia. Experimental results reveal that single Mn atoms were successfully anchored and highly dispersed on the surface of GDY, forming a catalyst with a determined chemical structure, facilitated charge transfer capability and maximized active sites. These special characteristics provide Mn<sub>SA</sub>/GDY with a high ammonia yield rate (up to 46.78 μg h<sup>-1</sup> mg<sub>cat</sub><sup>-1</sup>) and faradaic efficiency (39.83%), as well as cycling durability, which outperform those of most reported conventional atomic electrocatalysts and heterostructured ones. This work broadens the application scope of GDY-based atomic catalysts and demonstrates their applicability in the field of energy conversion.

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## Introduction

The ammonia industry has undergone tremendous development since its inception and has greatly changed the landscape of our planet; its impact on our daily life is enormous. Ammonia has been widely used to produce synthetic nitrogen fertilizers that are essential for crop harvesting.<sup>1–3</sup> Additionally, it has been used for producing numerous chemical compounds (*e.g.*, urea, ammonium nitrates, ammonium phosphates, and other nitrogen compounds) and industrial refrigerants. Today, more than 90% of ammonia is synthesized through the Haber–Bosch process,<sup>4</sup> which requires high-purity N<sub>2</sub> and H<sub>2</sub> and high temperature and pressure, thereby inducing harmful effects on the environment from a long-term perspective because of its fossil-fuel-dependent and energy-intensive features. Considering that the human-induced energy crisis and climate change have aroused international concerns, scientists have devoted tremendous efforts to explore innovative approaches for ammonia synthesis with the aim of reducing the energy consumption and CO<sub>2</sub> emissions from the current industrial ammonia synthesis process.<sup>5</sup>

Given that the cost of electricity generation has shown a decreasing trend annually, the electrocatalytic nitrogen

reduction reaction (ECNRR) powered by electricity generated from solar or wind sources has been considered as a promising and economically viable alternative approach for ammonia synthesis.<sup>4,6,7</sup> The ECNRR can be performed at ambient temperature and pressure in relatively small-scale, distributed, and on-site electrolysis cells. However, a critical research issue is the exploration of efficient electrocatalysts to break the inert molecular structure of N<sub>2</sub>, which has a high triple bond energy of 941 kJ mol<sup>-1</sup>, and facilitate subsequent delivery and transfer of protons and electrons.<sup>6</sup> Another major issue for ECNRR research is that the competitive hydrogen evolution reaction which limits the faradaic efficiency and leads to a low overall reaction rate should be hindered by electrocatalysts with high selectivity for nitrogen fixation and reduction. Fragile molecular-scale catalysts do not meet the requirements of long-term electrocatalysis, although they possess good NH<sub>3</sub> selectivity.<sup>4</sup> The challenges presented by fragile molecular-scale catalysts can be circumvented by using heterogeneous electrocatalysts.<sup>8</sup> Numerous studies have reported achieving great improvements in electrocatalytic activity, stability, and selectivity when using various heterogeneous electrocatalysts containing noble metals (Au, Ru, Pd, and Pt) and their alloys,<sup>9,10</sup> transition metals (Fe, Mo, Cu, and V) and their alloys,<sup>11</sup> nitrides,<sup>12</sup> oxides,<sup>11</sup> sulfides,<sup>13–17</sup> and metal-free catalysts.<sup>18–20</sup> Compared with the above-mentioned electrocatalysts, research on atomic catalysts with higher atomic utilization and cost-effectiveness appears to be scarce.

Graphdiyne (GDY), a novel two-dimensional carbon allotrope in which adjacent sp<sup>2</sup>-hybridized benzene rings are linked to each other *via* sp-hybridized butadiyne linkages, features a highly conjugated system with strong reduction ability, an

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intrinsic uniform porous structure and excellent mechanical properties.<sup>21–23</sup> It has been widely applied in various research fields, including catalysis, batteries, solar cells, and electrochemical actuators owing to its fascinating properties.<sup>24–29</sup> It is worth noting that the arbitrary angle of rotation ( $\pi/\pi^*$ ) perpendicular to the axis in the GDY structure enables it to chelate single metal atoms, thereby making it possible to use it as an ideal support for anchoring and preparing atom catalysts without any pre-treatment.<sup>24</sup> Many GDY-based atomic electrocatalysts have been successfully synthesized and demonstrated excellent activities towards various catalytic reactions.<sup>24,30–33</sup> Additionally, theoretical investigations of single and double transition metals (Fe, Mn, Nb, Cr, V, *etc.*) anchored on graphdiyne for nitrogen reduction reactions have recently demonstrated good catalytic activity and selectivity in nitrogen fixation under ambient conditions.<sup>34,35</sup> This outcome has inspired the current study with an aim of widening the application space of GDY-based atomic catalysts in the field of electrochemical ammonia synthesis. Therefore, this study demonstrates the first successful preparation of manganese atomic electrocatalysts anchored on the GDY substrate ( $\text{Mn}_{\text{SA}}/\text{GDY}$ ) and their promising properties in electrocatalytic nitrogen reduction reactions.

## Results and discussions

The freestanding  $\text{Mn}_{\text{SA}}/\text{GDY}$  catalytic electrode was prepared *via* a facile two-step strategy involving the initial synthesis of GDY nanosheets (GDY NSs) on nitric acid-treated carbon cloth *via* an acetylenic cross-coupling reaction. This step was followed by *in situ* anchoring of Mn atoms on GDY NSs through electrochemical reduction deposition at an overall mass loading of  $0.155 \text{ mg cm}^{-2}$  (Fig. 1a). Scanning electron microscopy (SEM) images (Fig. 1b–d) revealed the porous and ultrathin nanosheet morphology of the as-prepared GDY and  $\text{Mn}_{\text{SA}}/\text{GDY}$ , respectively. Inductively coupled plasma mass spectrometry (ICP-MS) analysis suggested that the mass loading of Mn atoms on GDY

NSs was 0.19 wt%. The contact angle measurement results (Fig. 1e) indicated the hydrophilicity of the  $\text{Mn}_{\text{SA}}/\text{GDY}$  surface, which could reduce the contact and transport resistance between the aqueous electrolyte and electrode, thereby facilitating the subsequent electrochemical reaction. Transmission electron microscopy (TEM) was conducted for the morphological characterization of porous  $\text{Mn}_{\text{SA}}/\text{GDY}$  (Fig. 2a). High-resolution transmission electron microscopy (HRTEM) images only showed the interplanar distance of GDY without any lattice fringe of the metal species (Fig. 2b and c). Subsequent high-angle annular dark-field scanning transmission electron microscopy (HAADF-STEM) images revealed the isolated and uniform dispersion of bright dots representing heavy-element atoms. This result confirmed the successful anchoring of Mn atoms on GDY (Fig. 2d–h). In addition, scanning TEM (STEM) and corresponding elemental mapping results indicated the presence of C and Mn elements (Fig. 2i).

Raman spectroscopy, X-ray diffraction (XRD) analysis and X-ray photoelectron spectroscopy (XPS) were also performed to investigate the chemical composition and structure of  $\text{Mn}_{\text{SA}}/\text{GDY}$ . The XRD spectra of both  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY (Fig. 3a) showed only the characteristic peaks of carbon. Raman spectra of pure GDY and  $\text{Mn}_{\text{SA}}/\text{GDY}$  (Fig. 3b) both suggested the typical D and G bands of carbon materials and peaks attributed to the vibration of conjugated diyne units ( $\text{C}\equiv\text{C}$ , peaks marked with “\*”) of GDY.<sup>21</sup> In contrast, the higher intensity ratio of D/G of  $\text{Mn}_{\text{SA}}/\text{GDY}$  than that of pure GDY implied its more defective property and the abundance of electrocatalytically active sites. XPS measurements were performed to explore the valence state and chemical bonding of the samples. The valence band spectra in Fig. 3d exhibited the variation trend of valence band maxima



Fig. 1 (a) Schematic representation of the synthetic process of  $\text{Mn}_{\text{SA}}/\text{GDY}$ . (b) SEM image of GDY. (c and d) SEM images of  $\text{Mn}_{\text{SA}}/\text{GDY}$  at different scale bars. (e) Contact angle measurements of  $\text{Mn}_{\text{SA}}/\text{GDY}$ .

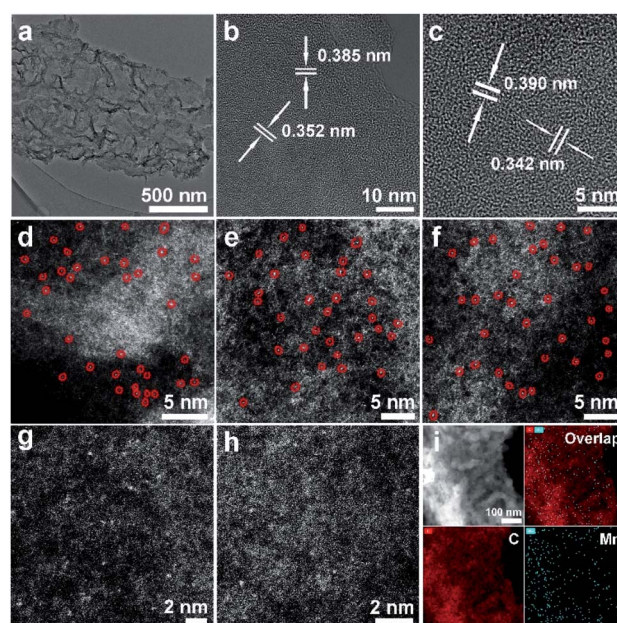


Fig. 2 (a) TEM image of  $\text{Mn}_{\text{SA}}/\text{GDY}$ . (b and c) HRTEM images of  $\text{Mn}_{\text{SA}}/\text{GDY}$ . (d–h) HAADF images of  $\text{Mn}_{\text{SA}}/\text{GDY}$  at different scale bars. (i) STEM image of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and corresponding elemental mapping results of  $\text{Mn}_{\text{SA}}/\text{GDY}$ .



Fig. 3 (a) XRD patterns of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY. (b) Raman spectra of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY. (c) XPS survey of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY. (d) Valence band spectra of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY. (e) Mn 2p XPS spectrum of  $\text{Mn}_{\text{SA}}/\text{GDY}$ . (f) C 1s XPS spectra of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY.

(VBM) of the samples where the VBM of  $\text{Mn}_{\text{SA}}/\text{GDY}$  shifted closer to the Fermi level, implying promoted conductivity after anchoring Mn atoms on GDY.<sup>30</sup> Fig. 3e shows the Mn 2p XPS spectrum of  $\text{Mn}_{\text{SA}}/\text{GDY}$  with two peaks located at 642.4 eV and 653.9 eV. These peaks correspond to the spin doublet of Mn 2p<sub>3/2</sub> and Mn 2p<sub>1/2</sub> with a separation of 11.5 eV. The fitting of four deconvoluted peaks in the high-resolution Mn 2p<sub>3/2</sub> XPS spectrum given in Fig. S1† indicated the presence of multiple Mn valences.<sup>36,37</sup> The C 1s XPS spectrum of  $\text{Mn}_{\text{SA}}/\text{GDY}$  (Fig. 3f) suggested a new sub-peak at 290.0 eV, which was attributed to the  $\pi-\pi^*$  transitions originating from the interactions between the GDY and heterogeneous metal species.<sup>21,22</sup> All these findings confirmed the synergistic effect between individually anchored Mn atoms and the GDY substrate which arises from the unique structural and electronic properties of GDY.

The electrocatalytic nitrogen reduction activities of  $\text{Mn}_{\text{SA}}/\text{GDY}$  were evaluated in neutral 0.1 M  $\text{Na}_2\text{SO}_4$  electrolyte using a gas purification set-up (Fig. 4a). The freshly prepared samples were kept in a glove box filled with Ar to prevent nitrogen compounds and other contaminants in the atmosphere from adsorbing on the surface. The feeding gas was carefully purified through a Cu trap to remove any possible contaminants containing  $\text{NO}_x$  or other nitrogen species before each test. The polarization curves obtained under  $\text{N}_2$  and Ar atmospheres (Fig. 4c) demonstrate the existence of the nitrogen reduction reaction. Fig. 4e and f exhibit the specific and areal ammonia yield ( $Y_{\text{NH}_3}$ ) and the corresponding faradaic efficiency (FE) of



Fig. 4 (a) Schematics of the cell configuration for the NRR under ambient conditions. (b) Calibration curve used for the quantification of produced ammonia. (c) Polarization curves recorded under  $\text{N}_2$  and Ar atmospheres. (d) UV-vis spectra of the electrolytes after the NRR at different potentials for 1 h under ambient conditions. (e) Specific ammonia yields and FEs of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and pure GDY at different potentials in  $\text{N}_2$ -saturated 0.1 M  $\text{Na}_2\text{SO}_4$  (bars present standard deviation). (f) Areal ammonia yields and FEs of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and pure GDY at different potentials in  $\text{N}_2$ -saturated 0.1 M  $\text{Na}_2\text{SO}_4$  (bars present standard deviation). (g)  $Y_{\text{NH}_3}$  of  $\text{Mn}_{\text{SA}}/\text{GDY}$  obtained at different potentials in Ar- and  $\text{N}_2$ -saturated 0.1 M  $\text{Na}_2\text{SO}_4$ , respectively (bars present standard deviation). (h) UV-vis spectra for the detection of  $\text{N}_2\text{H}_4$  on  $\text{Mn}_{\text{SA}}/\text{GDY}$  at applied potentials in  $\text{N}_2$ -saturated 0.1 M  $\text{Na}_2\text{SO}_4$ . (i) Durability test results of  $\text{Mn}_{\text{SA}}/\text{GDY}$  at  $-0.045$  V *versus* RHE in 0.1 M  $\text{Na}_2\text{SO}_4$ .

$\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY at different applied potentials from 0.155 V–0.245 V *versus* RHE, respectively. The corresponding UV-vis absorption curves and chronoamperometry curves are shown in Fig. 4d and S3a,† respectively. It was observed that  $\text{Mn}_{\text{SA}}/\text{GDY}$  displayed a higher  $Y_{\text{NH}_3}$  and FE than those of pure GDY. The highest specific  $Y_{\text{NH}_3}$  ( $45.35 \pm 1.38 \mu\text{g h}^{-1} \text{mg}_{\text{cat}}^{-1}$ , which could be converted into an areal  $Y_{\text{NH}_3}$  of  $(1.27 \pm 0.24) \times 10^{-10} \text{ mol s}^{-1} \text{cm}^{-2}$ ) and FE ( $38.08 \pm 2.26\%$ ) for  $\text{Mn}_{\text{SA}}/\text{GDY}$  were obtained at  $-0.045$  V *versus* RHE. These results outperform those of most atomic NRR electrocatalysts and even some nano-scale materials (Table S1†).<sup>38–46</sup> Electrochemical reduction reactions were performed under an Ar atmosphere to confirm the nitrogen source of the detected ammonia. No ammonia could be detected when Ar was applied as the feeding gas (Fig. 4g and S4†), verifying that the ammonia detected in the tests originated from the reduction of nitrogen rather than contaminants. UV-vis spectra ranging from 400 to 600 nm of the electrolytes after the NRR were measured to detect the possible  $\text{N}_2\text{H}_4$  by-product. As shown in Fig. 4h,  $\text{Mn}_{\text{SA}}/\text{GDY}$  exhibited 100% selectivity for ammonia production in the absence of the characteristic peak of  $\text{N}_2\text{H}_4$  at 455 nm at all applied potentials. The durability of the  $\text{Mn}_{\text{SA}}/\text{GDY}$  electrocatalyst was evaluated through a cycling test where the NRR was performed for 1 h at  $-0.045$  V *versus* RHE in each cycle. Fig. 4i indicates that the  $Y_{\text{NH}_3}$  and FE of  $\text{Mn}_{\text{SA}}/\text{GDY}$  could still be maintained after eight cycles, implying excellent stability. SEM and TEM images shown in Fig. S5† exhibited a well-retained morphological and chemical structure of  $\text{Mn}_{\text{SA}}/\text{GDY}$

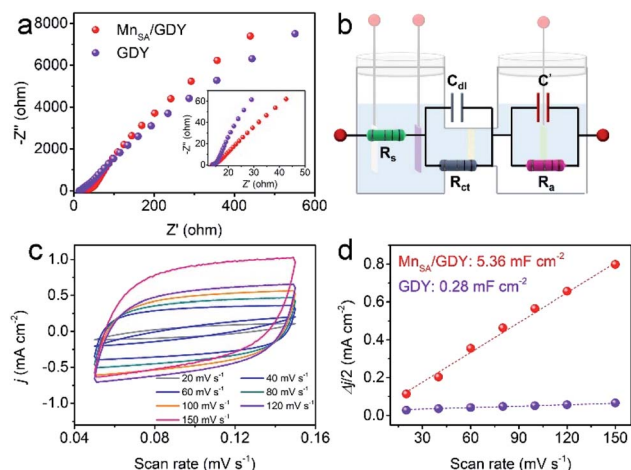


Fig. 5 (a) Electrochemical impedance spectra of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY (inset: simulated Nyquist plots of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY). (b) Corresponding fitted equivalent circuit model. (c) CV curves of  $\text{Mn}_{\text{SA}}/\text{GDY}$  at different scan rates. (d) Estimated  $C_{\text{dl}}$  values of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY.

after the durability test verifying the good structural stability of  $\text{Mn}_{\text{SA}}/\text{GDY}$  during electrochemical measurements.

Electrochemical impedance spectra (EIS) and electrochemically active surface area (ECSA) were measured to determine the intrinsic electrochemical properties of  $\text{Mn}_{\text{SA}}/\text{GDY}$ . The Nyquist plots of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY are shown in Fig. 5a, and fitted by an equivalent circuit model (Fig. 5b) comprising the solution resistance ( $R_s$ ) and charge transfer resistance ( $R_{\text{ct}}$ ). It was observed that both  $R_s$  and  $R_{\text{ct}}$  for  $\text{Mn}_{\text{SA}}/\text{GDY}$  decreased after the anchoring of Mn atoms on the GDY substrate, as compared to that of pristine GDY. This result indicates the more facilitated electron transfer capability of  $\text{Mn}_{\text{SA}}/\text{GDY}$  during the NRR. The electrochemical double layer capacitance ( $C_{\text{dl}}$ ) of  $\text{Mn}_{\text{SA}}/\text{GDY}$  and GDY was measured and is displayed in Fig. 5c and d. The higher  $C_{\text{dl}}$  of  $\text{Mn}_{\text{SA}}/\text{GDY}$  suggested its larger ECSA, which could afford more abundant active sites and thus result in its higher reaction efficiency.

We further conducted XPS measurements on  $\text{Mn}_{\text{SA}}/\text{GDY}$  after cycling stability tests (Fig. S7†). As shown in Fig. S7,† the N 1s XPS spectrum after cycling stability tests could be deconvoluted into two subpeaks corresponding to the N–H bond and adsorbed  $\text{N}_2$  respectively. This revealed the absorbance of  $\text{N}_2$  and ammonia products as well as the corresponding reaction intermediates on the surface of the catalyst. The lower binding energy of the C 1s XPS spectrum of  $\text{Mn}_{\text{SA}}/\text{GDY}$  after the reaction, as compared to that of as-prepared  $\text{Mn}_{\text{SA}}/\text{GDY}$ , imply charge transfer from Mn atoms to GDY during the electrocatalysis.

## Conclusions

In summary, we have successfully synthesized Mn atomic catalysts supported on a GDY substrate with an atomic mass loading of 0.19 wt%. This  $\text{Mn}_{\text{SA}}/\text{GDY}$  electrocatalyst exhibits a definite chemical structure, promoted charge transfer capability and abundant active sites. It demonstrates promising performance towards electrocatalytic nitrogen reduction to

ammonia, where the ammonia yield and FE could reach up to  $46.78 \mu\text{g h}^{-1} \text{mg}_{\text{cat}}^{-1}$  ( $1.54 \times 10^{-10} \text{ mol s}^{-1} \text{cm}^{-2}$ ) and 39.83%, respectively, and the activity could be retained well after eight cycling tests. These results exceed those of most reported conventional atomic electrocatalysts and heterostructured ones. This work further broadens the application space of GDY-based atomic electrocatalysts and proves their strong competitiveness in the field of catalysis.

## Conflicts of interest

There are no conflicts to declare.

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