


 Cite this: *RSC Adv.*, 2022, 12, 35556

MnO₂-based materials for supercapacitor electrodes: challenges, strategies and prospects

Juyin Liu, Jiali Bao, Xin Zhang, Yanfang Gao, * Yao Zhang, Ling Liu* and Zhenzhu Cao

Manganese dioxide (MnO₂) has always been the ideal electrode material for supercapacitors due to its non-toxic nature and high theoretical capacity (1370 F g⁻¹). Over the past few years, significant progress has been made in the development of high performance MnO₂-based electrode materials. This review summarizes recent research progress in experimental, simulation and theoretical studies for the modification of MnO₂-based electrode materials from different perspectives of morphology engineering, defect engineering and heterojunction engineering. Several main approaches to achieve enhanced electrochemical performance are summarized, respectively increasing the effective active site, intrinsic conductivity and structural stability. On this basis, the future problems and research directions of electrode materials are further envisaged, which provide theoretical guidance for the adequate design and synthesis of MnO₂-based electrode materials for use in supercapacitors.

 Received 21st October 2022
 Accepted 28th November 2022

DOI: 10.1039/d2ra06664e

rsc.li/rsc-advances

1. Introduction

In recent years, the rapid development of the global economy and the expanding industrialization of science and technology have led to the depletion of fossil energy sources and the production of large amounts of carbon dioxide emissions, which have a serious impact on the environment and human beings. In order to meet the demand for efficient energy storage systems, there is an urgent need to research and develop new electrochemical energy storage devices (EESDs) for energy conversion, which are suitable for different energy storage systems in multiple fields.¹ Supercapacitors (SCs), also known as electrochemical capacitors (ECs), have attracted wide attention as they are alternative to both conventional capacitors and secondary batteries with such features as high-power density, long cyclic stability, good rate capability and fast charge/discharge.^{2,3} SCs have unique advantages so that make them show good perspectives for future applications in new energy technologies, such as smart and wearable electrical products, wind pitch devices and grid access systems for power grids, *etc.*⁴⁻⁷ In addition, SCs can fill the gap between batteries and conventional capacitors.¹ They meet the demand for short-term acceleration in electric and hybrid vehicles by providing higher power density. The power density and energy density are represented by Ragone plot (Fig. 1).

Based on the energy storage mechanisms, SCs are generally defined in three major categories: electric double-layer

capacitors (EDLCs), Faraday pseudo-capacitors and hybrid capacitors which are made from a combination of the above two. In EDLCs, carbon-based materials are common electrode materials, which through the adsorption and desorption of ions

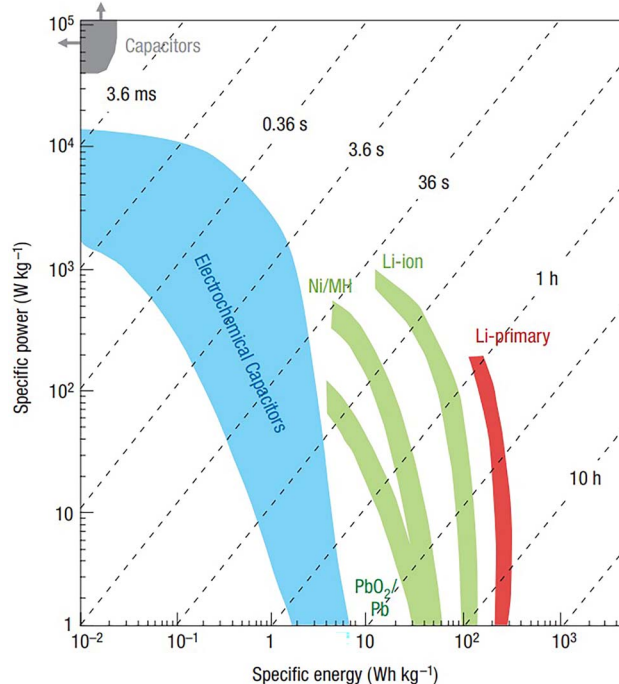


Fig. 1 Ragone plot of different electrochemical energy conversion systems. Reproduced with permission from ref. 1. Copyright 2008, Springer Nature.

School of Chemical Engineering, Inner Mongolia University of Technology, No. 49 Aimin Street, Xincheng District, Hohhot 010051, PR China. E-mail: yf_gao@imut.edu.cn; ll@imut.edu.cn



on the surface of the electrodes, complete the energy storage. Currently, many researchers have extensively studied carbon-based materials. Based on the traditional porous carbon materials, many new porous carbon materials have been prepared, which are able to achieve an increase in specific capacitance through the adjustment of the pore structure.^{8,9} However, due to the limited physical space on the surface, its obtained energy density is lower.^{10,11} In addition, the fast and reversible redox reacted inside the surface of electrodes in Faraday pseudo-capacitor that contribute to provide higher capacitance and increase energy density without impacting power density.¹² Nevertheless, the research on the matching degree of parameters between two electrodes (positive and negative) are not mature. Hence, it requires the development of good positive and negative electrode materials with high performance and compatibility.¹³ In comparison, hybrid capacitors can achieve higher capacity and energy storage capability by combining two energy storage mechanisms of EDLCs and Faraday pseudo-capacitors.¹⁴

In general, the electrode materials of SCs are divided into EDLCs materials and pseudo-capacitor materials. Compared with the EDLCs materials, because of the high specific capacitance of the pseudo-capacitor materials, it plays an important role in the process of obtaining good capacitance and energy density. So far, many electrode materials based on transition metal oxides have been explored, such as TiO₂,¹⁵ RuO₂,^{16,17} V₂O₅,¹⁸ Co₃O₄,¹⁹ Fe₃O₄,²⁰ MoO₃ (ref. 21 and 22) and MnO₂.^{23–26} Besides, spinel-based mixed transition metal oxides have also been investigated. Compared to the low electrical conductivity of oxides, spinel-based cobaltite has a variety of oxidation states and a higher electrical conductivity, which allows it to store more charge and thus achieve higher capacitive properties. These include FeCo₂O₄,^{27–29} ZnCo₂O₄,^{30–32} CuCo₂O₄,^{33–35} MnCo₂O₄,^{36–38} *etc.* MnO₂ has received great attention, because of its low price, low toxicity, high theoretical capacity (1370 F g⁻¹) and good environmental compatibility, *etc.*³⁹ But the electronic conductivity and structural stability of MnO₂ is poor, and there are some shortcomings such as the manganese (Mn) element is easily dissolved in the electrolyte, which limits the development and application of MnO₂-based electrode materials.⁴⁰ Therefore, researchers have usually worked on combining MnO₂ with highly conductive substrates (*e.g.*, metals or carbon-based materials) in order to enhance the whole performances. A series of MnO₂-based composites have been designed, including MnO₂/precious metal (gold, silver) composites,^{41–43} MnO₂/transition metal (Cu, Ni, Mn) composites,^{44–47} MnO₂/graphene composites,^{48–50} MnO₂/carbon nanotube (CNT) composites,^{51,52} MnO₂/porous carbon composites,^{53,54} *etc.* Up to now, MnO₂-based materials have got good experiment and calculation results in aspects of enhanced electrical conductivity and improved specific surface area, and they have also been reviewed extensively from different perspectives. Our primary intention is to summarize the various engineering strategies in MnO₂-based electrode materials through reference and guidance for the unique design and preparation of MnO₂-based electrodes by specifically.



Fig. 2 Different engineering strategies to improve electrochemical performance.

In this review, we intend to conduct an in-depth investigation and comprehensive overview of the crystal structures and energy storage mechanisms of MnO₂-based electrode materials firstly. In addition, the problems which limit the application of MnO₂-based electrode materials are proposed, and several examples are combined to summarize and compare different engineering strategies that can improve the performance of MnO₂-based SCs (Fig. 2). Through in-depth research on improving the electrical conductivity and structural stability of MnO₂, and providing effective active sites for electrolyte cations. Finally, we give the overall summary of MnO₂-based electrode materials as well as the future outlook.

2. Types of MnO₂ crystal structures

In general, MnO₂ consists of [MnO₆] octahedral unit, which is composed of one Mn atom and six oxygen (O) atoms. These [MnO₆] octahedral units are connected to each other in different ways through shared angles and shared chains and so it shows a variety of tunnel and chain structures, which correspond to different crystal structures.⁵⁵ Among them, the crystal form of MnO₂ are mainly included: α -MnO₂, β -MnO₂, δ -MnO₂, γ -MnO₂, ϵ -MnO₂ and λ -MnO₂. The different crystal structures are shown in Fig. 3. According to the space structures, MnO₂ can be divided into three categories: one-dimensional (1D) tunnel structures, two-dimensional (2D) layered structures and three-dimensional (3D) network structures.⁵⁶

The different MnO₂ have different structures and space groups, as shown in Table 1. α -MnO₂, β -MnO₂, γ -MnO₂ belong to the 1D tunnel structures where the lattice is all made up of shared chains and shared angles.^{57,58} α -MnO₂ is a tetragonal crystal system, which belongs to the typical hollandite. It is crosslinked by the [MnO₆] octahedron shared edge double



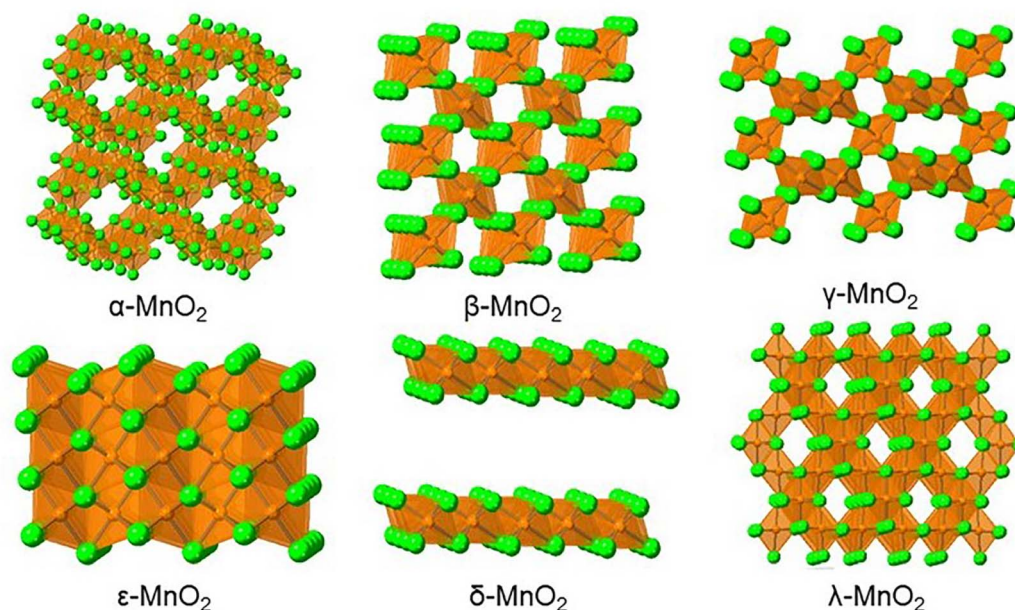


Fig. 3 Schematic diagram of the crystal structure of α -, β -, γ -, ϵ -, δ - and λ -MnO₂.

chain and forms 1D (2×2) tunnel structure through corner connections. The tunnel size of α -MnO₂ is 4.6 Å where most of the cations (e.g. K⁺, Li⁺, Na⁺, Mg²⁺, Ca²⁺, etc.) can be allowed to pass through for its larger tunnel. β -MnO₂ is also a tetragonal crystal system with a rutile structure. The Mn atom is used as the centre of the [MnO₆] octahedron, and O atoms occupy its hexagon to arrange into a single-chain angular close-packed structure. The 1D (1×1) tunnel structure is formed by means of shared edge connections where each chain is connected by four similar chains. The narrow tunnel of β -MnO₂ is about 1.89 Å. It can only accommodate smaller ions, which is not conducive to the diffusion of ions. γ -MnO₂ is a hexagonal crystal system with a hexagonal close-packed structure.⁵⁹ γ -MnO₂ is the result of irregular alternating symbiosis of pyrolusite (1×1 type) and rhodochrosite (1×2 type). ϵ -MnO₂ has a structure that is comparable to γ -MnO₂, and it belongs to the hexagonal crystal system as well. The former is only connected by shared planes, which leads to a disordered appearance of its lattice and the formation of an irregular tunnel structure. Mn⁴⁺ is distributed in more than half of the [MnO₆] octahedral gap.⁶⁰ Both δ -MnO₂ and λ -MnO₂ are formed by using only the shared edges of

[MnO₆] octahedron. δ -MnO₂ belongs to the monoclinic crystal system, which has 2D layered structure formed by the common edges of the [MnO₆] octahedron. It has a large interlayer distance of about 7 Å, which can allow a large number of water molecules, metal cations and other substances to pass through. Consequently, δ -MnO₂ is beneficial to the migration of ions/substances. λ -MnO₂ is a spinel structure, which is similar to γ -MnO₂. After the irregular alternating growth of pyrolusite and rhodochrosite, it forms a 3D (1×1 type) network structure, which promotes charge transfer.

3. Energy storage mechanism of MnO₂

Electrochemical behaviors of MnO₂-based electrode materials are divided into two types according to their energy storage mechanisms: EDLCs materials and Faraday pseudo-capacitance materials. The properties are determined by the cations in the electrolyte which lead to different energy storage mechanisms of MnO₂-based electrode materials. For example, it usually shows pseudo-capacitive behavior in electrochemical

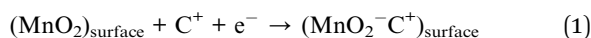
Table 1 Multiple parameters of MnO₂ crystal structures

Crystalline forms	Structure types	Space groups	Tunnels ($n \times m$)	Size (Å)	Ref.
α -MnO ₂	Hollandite	$I4/m$	(2×2)	4.6	61
β -MnO ₂	Pyrolusite	$P4_2/mnm$	(1×1)	1.89	62
γ -MnO ₂	Distorted boehmite	—	(1×1)/(1×2)	1.89/2.3	63
δ -MnO ₂	Birnessite	$C2/m$	($1 \times \infty$)	7.0	64
ϵ -MnO ₂	—	$P6_3/mmc$	(1×1)/(1×2)	3D	65
λ -MnO ₂	Spinel	$Fd3m$	(1×1) interconnected 3D networks	3D	66



performance tests when the charge–discharge process occurs in alkaline electrolytes or neutral electrolytes containing alkali metal ions (Li^+ , Na^+ , K^+ , *etc.*).^{67,68}

In the energy storage mechanism of EDLCs, MnO_2 -based electrodes exhibit adsorption/desorption activities over the material surface. After immersing the electrode in electrolyte, a tight charge layer is formed around the electrode surface by applying a certain voltage between the positive and negative electrodes, which promotes the separation of charges.^{69–71} But in most cases, the Faraday pseudo-capacitor energy storage mechanism is commonly used to analyze. Harnessing the variable valence states of pseudocapacitive materials to store charges. According to the differences in the thickness of MnO_2 -based electrode materials, it is further divided into two kinds of charge storage mechanisms for Faraday pseudo-capacitor. On the nanoscale MnO_2 -based electrode layers for surface charge storage. The reversible faradaic reaction on/near the electrode surface can be accomplished by adsorption/desorption of ions which from electrolyte. The process is quite similar to the physical adsorption/desorption in an EDLC, and the difference is that electrochemical reactions take place in fact. It is usually expressed by eqn (1):



In the formula, $\text{C}^+ = \text{Li}^+$, Na^+ , K^+ , *etc.* Since the electrolyte cations have opposite charges to the O atoms in the $[\text{MnO}_6]$ octahedron, and they are attracted to the O atoms and become electrically neutral during the charge/discharge. The additional generated charges are transferred to the nearby Mn atom, which results in the change of Mn^{4+} to Mn^{3+} .⁷² Furthermore, electrolyte cations or protons can also be used for charge storage through intercalation/deintercalation in the bulk phase of MnO_2 -based electrode materials.^{73,74} Eqn (2) is expressed as follows:



In the formula, $\text{C}^+ = \text{Li}^+$, Na^+ , K^+ , *etc.* Generally, the bulk pseudo-capacitive reaction is limited to the subsurface layer of MnO_2 , with a thickness of about 420 nm according to Brousse's study.⁷⁵ Due to the thickness of the bulk MnO_2 -based electrode materials, the resistance to the diffusion of electrolyte cations or protons in its interior is larger. It affects the rate of ion diffusion and charge transfer, and its conductivity is poor as well. Thus, it is further illustrated that the thickness of the MnO_2 -based electrode materials is crucial to the capacitance.

4. Challenges of MnO_2 -based electrode materials

4.1 Effect of active sites on capacitance

The accommodating level of MnO_2 -based electrode materials for protons or cations directly determines its charge storage capacity, which affects its electrochemical performance. Generally, adjusting the size of the crystal to increase the specific surface areas is a way to obtain higher specific

capacitance. For various crystal forms, the capacitance values exhibited by MnO_2 -based electrode materials are significantly different, which mainly depends on the size of spaces such as tunneling or interlayer spacing in the crystal structure of different dimensions. It is found that the capacitance value decreases for different crystal forms of MnO_2 , which following is the order: $\alpha > \delta > \gamma > \lambda > \beta$.^{55,76} However, due to the small proportion of effective Mn atom active centers in the crystal structure, which can attract fewer negative charges. With the increase of crystal structures and particle sizes of MnO_2 of a specific phase, the electrochemical performance is reduced. Therefore, increasing the specific surface areas and exposing more effective active sites have become a necessary way to improve electrochemical performance.

4.2 Influence of intrinsic conductivity on electrochemical properties

On the other hand, electrical conductivity is also one of the important factors affecting the electrochemical performance of electrode materials. Although MnO_2 -based electrode materials have fast and reversible redox reactions. Actually, the capacitance results of pure phase MnO_2 is only 14–43% of the theoretical capacity can be achieved. This phenomenon may be due to the electrons/ions transport *via* the MnO_2 electrode|electrolyte interface is relatively slow. Besides, the internal resistance is relatively large during the process of charging and discharging, which seriously reduces its rate capability.⁷⁷ The conductivity of MnO_2 -based electrode materials is determined by its own inherent structural properties, which make the actual capacity far from the theoretical capacity. However, the semiconductor property can determine the internal structural properties of MnO_2 , and its band gap is 0.25 eV. Therefore, to obtain better electrochemical performance, MnO_2 must reduce its band gap to favor metal characteristics. Hence, the key to improving the electrical conductivity of MnO_2 -based electrode materials lies in how to improve the limited electron transfer on the surface and the interior, that is, change its electronic properties.

4.3 Effect of crystal structure on cyclic stability

Finally, the structural stability as another important factor in evaluating the performance of electrode materials to determine the cyclic stability of the materials. In all electrode materials which involving MnO_2 , the unavoidable disproportionation reaction of Mn element causes the problem that Mn always dissolves slowly in the electrolyte.⁷⁸ The role of Mn atoms is to absorb excess charges, because the reduction of Mn leads to a decrease in the storage capacity of charges and affects the final cyclic stability of the materials. The size of the ions in the electrolyte and the spatial structure of the MnO_2 are not match. The crystal structure of MnO_2 expands and collapses when protons or cations are inserted or extracted, respectively. Consequently, it is urgent to improve the structural stability of MnO_2 .



5. Engineering strategies to improve electrochemical performance

In view of the above challenges faced by MnO₂-based electrode materials, researchers have proposed methods such as morphology control, introduction of defects or doping of various elements, and construction of heterostructures to improve certain properties of MnO₂ to achieve the improvement of its overall electrochemical performance.

5.1 Morphology engineering

Morphology engineering is a common strategy to regulate the specific surface area of electrode materials. Larger specific surface area can be achieved by controlling the morphology of electrode materials and obtained to promote its electrochemical performance. Since the size of MnO₂ grains is inversely proportional to its electrochemical performance in a certain range, and it is of great significance to synthesize nanoscale MnO₂-based electrode materials. In the past period of time, many MnO₂-based nanomaterials with different morphologies have been studied, and they have been designed and prepared by various synthetic methods. After a large number of experiments and literature review, we have summarized the existing morphologies of MnO₂-based materials. Various forms of nanomaterials exist, including zero-dimensional (0D), 1D, 2D and 3D nanomaterials according to the spatial dimension, such as nanospheres, nanowires, nanorods, nanotubes, nanofibers, nanosheets, nanobelts, nano-urchins and nanoflowers (Fig. 4).

Well-constructed nanostructures can effectively shorten the diffusion distance of ions to the electrode surface and the transport path of electrons. For example, Singu *et al.*⁷⁹ prepared MnO₂ films grown on stainless steel substrates by using the

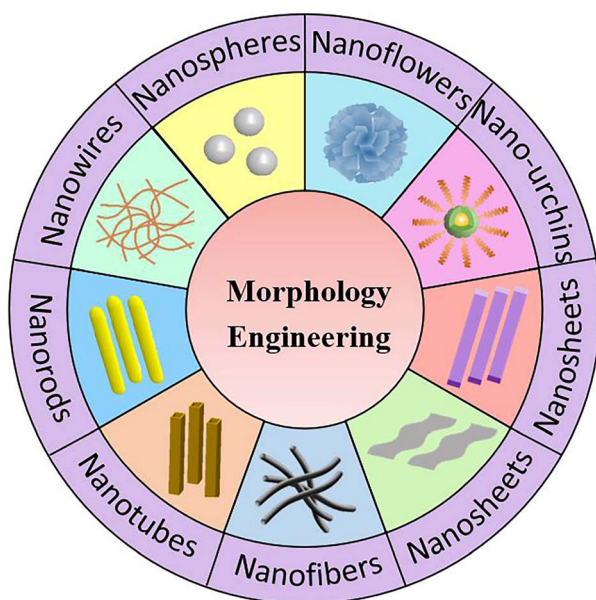


Fig. 4 Morphology engineering and structure of MnO₂-based materials.

sequential ionic layer adsorption and reaction method (SILAR). Three samples were synthesized according to the cycle numbers of SILAR. With the increase of cycle numbers, the diameter of the MnO₂ structure enlarged, which resulted in the change of nanoparticles into nanospheres (Fig. 5a–c). A thin film consisting of MnO₂ nanospheres in the form of α -MnO₂ and λ -MnO₂ co-existing was created eventually. The curves of the samples showed a typical triangular shape after charging and discharging tests when the diameter of MnO₂ nanospheres was 5–20 nm (Fig. 5d and e). It obtained specific capacitance up to 262 F g⁻¹ and better capacitance retention in Na₂SO₄ solution (Fig. 5f). It could be concluded that nanoparticles with certain porosity were favorable for the diffusion of electrolyte ions. Tang *et al.*⁸⁰ obtained MnO₂ nanomaterials with different morphologies by changing the hydrothermal reaction times under the mild conditions. The results indicated that the morphology of MnO₂ nanomaterials depended on the hydrothermal time. When the time of hydrothermal heating was increased, the MnO₂ crystallinity increased gradually. And this leads to the morphology changed from the initial nanowhisker spherical to α -MnO₂ nanorods. It led to an increase in the pore volume of α -MnO₂, and the specific surface area was also changed from 89 m² g⁻¹ to 119 m² g⁻¹. The better specific capacitance (152 F g⁻¹ at 2.5 mA) was displayed when the hydrothermal time was 6 h, which had good cyclic stability. In a word, changing the hydrothermal time is beneficial to change the nanoparticles to a single nanorod structure with higher specific surface areas, which contributed to better entry of electrolyte ions to obtain higher capacitance. Yin *et al.*⁸¹ reported a facile hydrothermal route for the synthesis of uniform and ultralong α -MnO₂ nanowires in a controllable redox reaction. Based on the nanomaterials obtained, it has many less porous with a size of \sim 3.8 nm. The current density of 1 A g⁻¹ showed a specific capacitance of 180 F g⁻¹, and the capacitance retention was above 78% after 2000 cycles. α -MnO₂ nanowires with large specific surface area and small porosity shorten the diffusion paths of electrons and ions, thereby enhancing the electrochemical performance of SCs.

Xu and co-workers⁸² synthesize highly loose mesoporous clusters of hollow spheres and sea urchin-structured α -MnO₂ nanomaterials in a simple hydrothermal method through the “Ostwald ripening process”. Among them, the shell of the hollow sea urchin was composed of densely arranged nanorods with consistent diameter and length, and the hollow sphere was composed of nanosheets. Due to its high specific surface area, it provided good cycling capability and high rate capacity. Xie *et al.*⁸³ developed a self-supporting material which composed of interconnected δ -MnO₂ nanosheets on graphite paper (GP). It demonstrated excellent specific capacitance (446 F g⁻¹ at 1 A g⁻¹) and outstanding cyclic stability under a wide voltage window of 1.3 V. Owing to the interaction which between δ -MnO₂ and the open nanosheet structures interconnected with the GP substrates, the effective contact of the electrolyte was promoted. It increased the specific surface area and improved ion diffusion and charge transfer.

We also investigated the morphology of pure phase MnO₂ materials. MnO₂ with different morphologies was obtained by





Fig. 5 Morphology of (a) sample 1, (b) sample 2 and (c) sample 3 prepared by SILAR method. (d) Comparative cyclic voltammetry (CV) of the three samples (e) The galvanostatic charge/discharge (GCD) curves of sample 2. (f) Specific capacitance of three samples with effect on scanning rate. Reproduced with permission from ref. 79. Copyright 2016, Elsevier B.V.

changing the hydrothermal temperature and reaction time under the condition of using the same precursor. For example, it was found that the obtained one-dimensional MnO_2 micro-particles showed a “moss”-like uniform distribution (Fig. 6a). However, the MnO_2 samples existed in the form of “dendritic” nanorods and packed together to form clusters when both the hydrothermal temperature and time were increased (Fig. 6b and c). And its electrochemical performance was better with a specific capacitance of 257 F g^{-1} at 1 A g^{-1} when hydrothermally reacted at 160°C for 12 h (Fig. 6d and e).

Generally speaking, the pseudo-capacitance properties of MnO_2 lead to a fast and reversible redox reaction, mainly

between Mn(IV) and Mn(III) . It enters protons and electrons into the lattice during the charging process, which means that the electrons released from the negative electrode enter the lattice of MnO_2 and reduce Mn^{4+} to Mn^{3+} . At the same time, the protons in the solution combine with O^{2-} in the lattice to become OH^- . Thus, for every electron and proton gained, one MnO_2 molecule is converted to MnOOH . The positive MnOOH releases H^+ to convert MnO_2 . Athouël *et al.*⁸⁴ doped Mg in layered MnO_2 and investigated in depth the effect of cycling processes on the crystal structure of MnO_2 by voltammetric behavior and XRD analysis. The results show that Mg as a dopant promoted the charge transfer process and that the

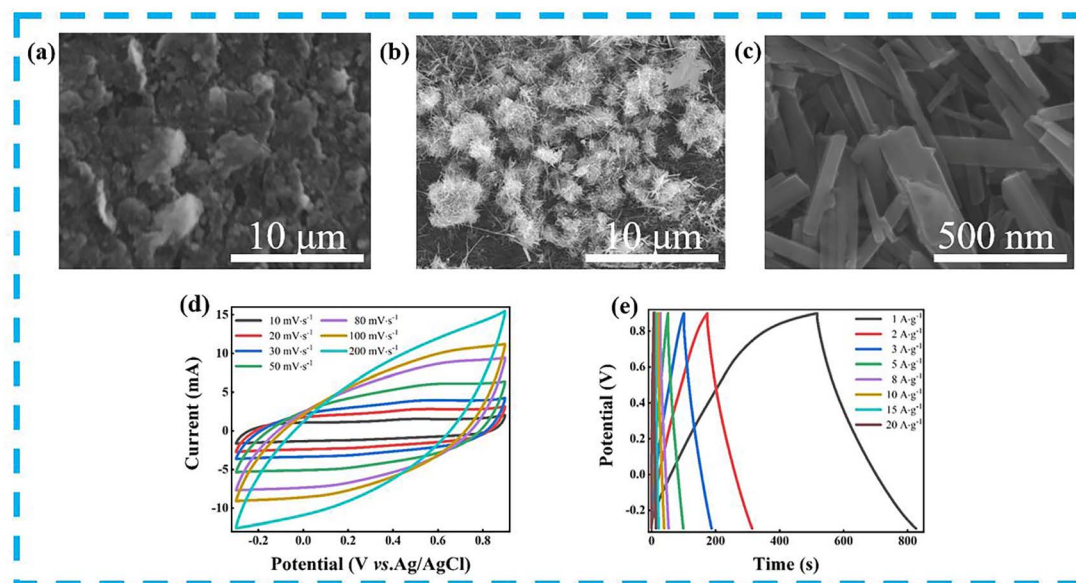


Fig. 6 SEM images of (a) sample 1 ($10 \mu\text{m}$), (b) sample 2 ($10 \mu\text{m}$) and (c) sample 2 (500 nm). (d and e) CV and GCD curves of sample 2.



charging and discharging process does not change the monoclinic structure of a classic Na-birnessite.⁸⁵ Therefore, the cycling process does not affect the crystal structure of MnO₂. The main reason for changing its structure is the size of the cations during the intercalation/deintercalation process. When the diameter of the embedded cation is smaller than its structural tunnel, the structure does not change before and after cycling. Suppose the diameter of the cation is larger than the structural tunnel. Its structure is disrupted during the intercalation of the cation into the crystal structure, resulting in an extremely rapid decrease in capacitance. In addition, the morphology of MnO₂ is more macroscopic than the valence and structure; therefore, the cycling process does not affect it significantly.

The MnO₂ structure changes fundamentally with increasing spatial dimensionality. A variety of morphologies are designed and composed, and the advantages of dimension control can be used to improve the shortcomings of one or several aspects of MnO₂-based electrode materials. Changing the morphology of electrode material, the higher porosity can be obtained to facilitate the diffusion of ions and electrons, or a larger specific surface area can be designed to expose more abundant active centers. This leads to further improvement of the electrochemical performances of MnO₂ nanomaterials in SCs.

5.2 Defect engineering

Due to the inherent internal limitations of the material, it is sometimes impossible to solve the existing problems with simple external modifications. Defect engineering is considered to be able to adjust the electronic structure of materials effectively, and the introduction of defects in MnO₂-based materials can provide them with new electronic, magnetic and optical properties.^{86,87} According to the mechanism of defect formation can be grouped into interstitial defects, substitutional defects, atomic vacancies and high-dimensional defects.⁸⁸ This section focuses on common defect engineering strategies for MnO₂-based materials: (1) oxygen vacancies, (2) cation vacancies, (3) cation doping and (4) anion doping.

5.2.1 Oxygen vacancy. Oxygen vacancy (V_O) is an important defect engineering strategy that enhances electrical conductivity by using V_O as an electron donor, while utilizing the modulation of the active center state to activate the intrinsically low-valent metal center, which in turn induces the formation of a new band gap. Zhu *et al.*⁸⁹ developed a method based on solubility contrast. Taking advantage of the different solubility of inorganic salts in different solvents, δ-MnO₂ with controllable nanostructures and V_O was obtained through a simple coprecipitation process. The δ-MnO₂ had a valence state mixed with the Mn³⁺-O-Mn⁴⁺ electron path after the addition of V_O, which contributed to improve electrical conductivity and electrochemical performance significantly. Through density functional theory (DFT) simulation calculations, they were found that the introduction of V_O can reduce the bond length of the Mn-O which was to effectively increase the Mn-O bond strength. After the introduction of V_O, three levels which closed to the Fermi level were generated. Compared with pristine

MnO₂, the electrical conductivity was increased (Fig. 7a and b). Liu *et al.*¹³ proposed a novel mechanism by utilizing self-triggered V_O for redox reactions of multiple Mn ions. An improvement in the electronic structure of MnO₂/Mn₂O₃ is achieved. It led to an abundance of V_O in the active structure inducing the new electronic states and electron redistribution between Mn and O, which resulted in a more than two-fold increase in the total charge storage. While promoting the redox reaction of multiple Mn ions (Mn²⁺, Mn³⁺, Mn⁴⁺). The Gibbs free energy of Na⁺ adsorption in the pristine MnO₂ and V_O-MnO₂ states and the density of states (DOS) of V_O-MnO₂ were calculated by DFT simulations. The simulation results showed that the adsorption energy of Na⁺ on the V_O-MnO₂ surface was -3.14 eV lower than that of -2.64 eV on the intact MnO₂ surface, which further confirmed that the presence of V_O was more favorable for the transfer of Na⁺. It could be imparted with a new electronic state adjacent to the Fermi level to introduce V_O in MnO₂. The electronic states could increase the carrier concentration, which improved conductivity and facilitated the transition of electrons from the valence to the conduction band by acting as a springboard. The difference in electron density for Na⁺ adsorption was compared between MnO₂ without V_O and with V_O. Na⁺ tended to absorb O with significant electrons, which indicated that O was the active site for Na adsorption. In particular, V_O weakened the chemical bond formed by Na and O on the surface, which promoted the adsorption/desorption of Na⁺, and also provided additional adsorption sites for Na⁺. It stabilized the layer structure by promoting the oxidation of Mn ions (Fig. 7c).

Recently, Zhang *et al.*⁹⁰ successfully introduced abundant V_O into the bulk MnO₂ phase by complex-induced chemical precipitation. MnO₂ with abundant bulk V_O exhibited fast charge transfer kinetics (Fig. 7d). The electronic properties of MnO₂ were investigated *via* DFT calculations, and the results showed that V_O could induce a reduction significantly in the band gap to enhance the conductivity. The concentration of V_O was proportional to the conductivity (Fig. 7e). Charge density difference calculations evidenced that electrons around O and Mn atoms were accumulated and depleted, respectively (Fig. 7f). Bader charge analysis showed that with increasing concentration of V_O, there was a greater electron transfer and a higher electron delocalization, which resulted in positive and negative region at V_O and around O, respectively. The adsorption energy of Na⁺ belonged to V_O-rich MnO₂ in the electrolyte was stronger than that of MnO₂ without V_O, which proved the presence of V_O was beneficial to the storage of Na⁺. The barrier of 4O_{cut}-MnO₂ (0.10 eV) was lower than that of MnO₂ (0.29 eV) and 2O_{cut}-MnO₂ (0.29 eV), which demonstrated that 4O_{cut}-MnO₂ had a fast Na⁺ transfer capability (Fig. 7g). At the same time, other reports had pointed out that the Mn 3d orbital state in MnO₂ with V_O was closer to the Fermi level and the state was more defined than that of MnO₂ without V_O. It was shown that V_O promoted the delocalized movement of nearby electrons to the vicinity of the low-coordinated Mn atom, causing the rapid transfer of the delocalized electrons into the conduction band and improving the overall conductivity of the system.⁹¹ Thus, abundant bulk V_O could increase the conductivity significantly by reducing the



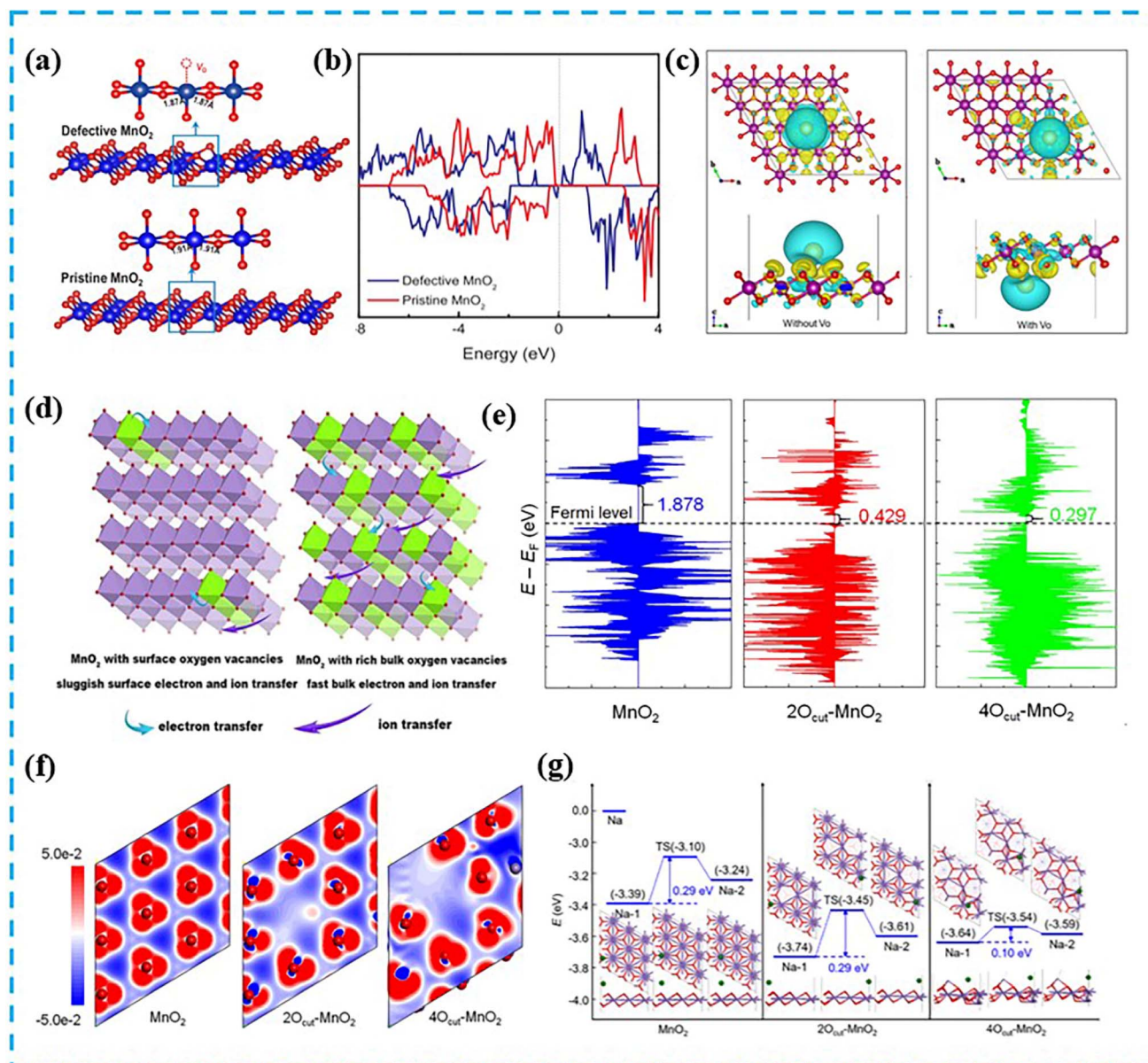


Fig. 7 (a) Constructed models of defective and perfect δ -MnO₂ surfaces. (b) Projected DOS of δ -MnO₂. Reproduced with permission from ref. 89. Copyright 2019, Elsevier Ltd. (c) Difference in electron density of MnO₂ without V_O and with V_O, blue and yellow represent charge consumption and accumulation (the Mn, O and Na atoms in purple, red and gold, respectively). Reproduced with permission from ref. 13. Copyright 2022, Elsevier B.V. (d) Charge transfer kinetics on MnO₂ with surface containing V_O and abundant bulk V_O. (e) DOS of MnO₂, 2O_{cut}-MnO₂ and 4O_{cut}-MnO₂ models. (f) Charge density difference between three models, with red and blue regions representing the accumulation and depletion of charge density, respectively. (g) Diagram of the energy distribution of Na⁺ diffusion on three models and the related structures of the intermediate, with the potential obstacles in blue (green, purple and red represented Na, Mn and O, respectively). Reproduced with permission from ref. 90. Copyright 2021, Elsevier B.V.

band gap and increasing the electron delocalization. Furthermore, a strong local electric field was formed around V_O, which could accelerate bulk ion transfer with a low diffusion energy barrier. In order to enhance the energy storage capacity of MnO₂ effectively, Jing *et al.*⁹² synthesized microsphere electrode materials using a typical hydrothermal-electrostatic self-assembly method combined with heat treatment. Based on *in situ* characterization techniques and DFT calculations, the pseudocapacitive behavior of K⁺ intercalation during discharge

was revealed. As the amount of K⁺ intercalation increased, the length of the Mn–O bond near K⁺ in the [MnO₆] octahedral structure became longer, which led to a decrease in its stability. As a result, the Mn–O bond lost O and generated V_O, so that the phase transition occurred and layered MnO₂ was transformed into a stable spinel-type Mn₃O₄.

Our research group had also successfully modified MnO₂ materials by means of V_O. After the method of low temperature calcination, MnO₂ containing a certain amount of V_O was



obtained by calcining at 300 °C for 4 h. In addition, the structure of CeO₂-coated MnO₂ was also formed through the additional introduction of Ce³⁺ ions. During this process, the morphology changed from nanorods to clusters formed by aggregation of nanospheres, and V_O also increased by 9.36% compared with the previous MnO₂. Therefore, the charge transfer resistance of the finally obtained material was reduced, and the electrochemical performance of the corresponding material was improved.

5.2.2 Cation vacancy. Cation vacancy can provide additional cation intercalation sites to effectively boost the charge storage capacity of transition metal oxides.⁹³ Thus, Mn vacancy is the only and most important cation vacancy strategy in MnO₂-based electrode materials. Wang *et al.*⁹⁴ synthesized monolayer δ -MnO₂ nanosheets with Mn vacancies by exfoliating their bulk counterparts in solution. By controlling the temperature of the prepared defect layered manganese oxide precursor, the concentration of Mn vacancies could be easily adjusted. The results showed that the lower the temperature, the more Mn vacancies were obtained. After calculations of DFT, the authors explored the implications of different types, concentrations and distributions of vacancies on the electronic structure of monolayer δ -MnO₂ nanosheets. As shown in Fig. 8a, when Mn vacancies were introduced into monolayer δ -MnO₂, it was transformed from semiconducting to semi-metallic properties due to the hybridization between the adjacent O 2p and Mn 3d orbitals around the Mn vacancies. With the increase of Mn vacancy concentration, the energy gap decreased gradually, but its semimetallic behavior was not affected. However, neither the introduction of no vacancy nor the introduction of single V_O could change the electronic properties of δ -MnO₂. In addition to using redox heat treatment to control the content of cationic

defects, it could be achieved by balancing the pH to control oxides in the solution as well. Gao *et al.*⁹⁵ applied the above method to strip and re-assemble δ -MnO₂ nanosheets into 3D macroporous pseudo-capacitor electrodes with controlled Mn point defect concentration and Mn³⁺/Mn⁴⁺ ratio. The defects provided two possible intercalation sites for Mn vacancies and the undercoordinated [MnO₆] surface octahedron on the opposite side of the nanosheets (Fig. 8b). The decrease in pH led to an increase in the Mn³⁺/Mn⁴⁺ ratio, which not only increased the concentration of Mn vacancies, but also enhanced the electrical conductivity and charge transfer efficiency due to the participation of high concentrations of Mn³⁺ in polaron hopping conduction (Fig. 8c). Meanwhile, results of electrochemical performance showed that the specific capacitance was improved from 200 F g⁻¹ to more than 300 F g⁻¹, and the charge transfer resistance was also significantly reduced. Therefore, the Mn vacancies in δ -MnO₂ nanosheets increased Na⁺ intercalation by providing new and low-energy intercalation sites, which enhanced cyclic stability.

5.2.3 Cationic doping. Dopants with similar ion sizes and high solubility properties can provide MnO₂-based materials with a variety of doping element choices, thereby tailoring the electrochemical properties of MnO₂-based materials according to practical applications. Doping cations (such as V, Li, Na, K, Co, Ag, La, Zn, Cu, Gd, and H) into the lattice of bulk MnO₂-based materials can release the electrochemical activity of MnO₂.⁸⁸ Cationic doping can be divided into interstitial doping and substitution doping generally. Tseng *et al.*⁹⁶ introduced Li⁺, Na⁺ and K⁺ with different doping concentrations into α -MnO₂ nanotubes respectively, and discussed the effect of doping of several metal cations on the crystal structure of α -MnO₂. According to DFT calculations, Mn³⁺ was formed by the electron



Fig. 8 (a) The charge density distribution and density of states of MnO₂ plates without Mn vacancy and with one Mn vacancy in 4 × 4. Reproduced with permission from ref. 94. Copyright 2015, Wiley-VCH. (b) Empirically estimated Mn vacancy concentration (circles) using Gaussian peak and linear baseline. (c) Average oxidation state of Mn from K-edge energies for δ -MnO₂ samples and standards obtained in different pH values. Reproduced with permission from ref. 95. Copyright 2017, Springer Nature.



transfer of K^+ doping and caused the Fermi level to overlap with the conduction band, which showed metallic behavior. The geometry did not change significantly at low K^+ doping concentration, it led to lattice expansion while at high doping concentration. The degree of expansion was proportional to the doping concentration, and the peak of the (211) plane in XRD shifted to lower 2θ values (Fig. 9a). Therefore, the doping concentration was so high that some K^+ may enter the interstitial sites, resulting in a disordered state of the structure. The significant difference in p density of state (PDOS) demonstrated that the incorporation of K^+ in the tunnel led to an enlarged d -spacing (Fig. 9b). And because the atomic radii of Li^+ and Na^+ were smaller than K^+ , the resulting lattice spacing was smaller. The energy splitting was asymmetric due to lattice distortion, which in turn broke the Mn symmetry between the $[MnO_6]$ octahedron. Yao *et al.*⁹⁷ proposed a strategy for obtaining

structural distortion of MnO_2 by doping Ni, while achieving improved electrochemical performance through fast electron and ion transfer kinetics. Ni- MnO_2 exhibited structurally distorted unit cells and expanded unit cell volume (Fig. 9c). According to the changes of Mn valence state and coordination number, it was shown that Ni doping in MnO_2 led to structural distortion in the crystal framework due to the changed of Mn-O bonds and Mn centers. At the same time, the electronic state modulation effect of electron delocalization and narrow band gap were used to enhance the conductivity. Therefore, the substitution of Mn in the octahedron by Ni not only effectively suppressed the Jahn-Teller distortion of Mn^{3+} in the structure to stabilize the MnO_2 structure and obtained excellent cyclic stability, but also provided a larger lattice parameter. It helped to lower the potential barrier of Na^+ to facilitate ion diffusion and accelerate the ion transport kinetics (Fig. 9d). In another

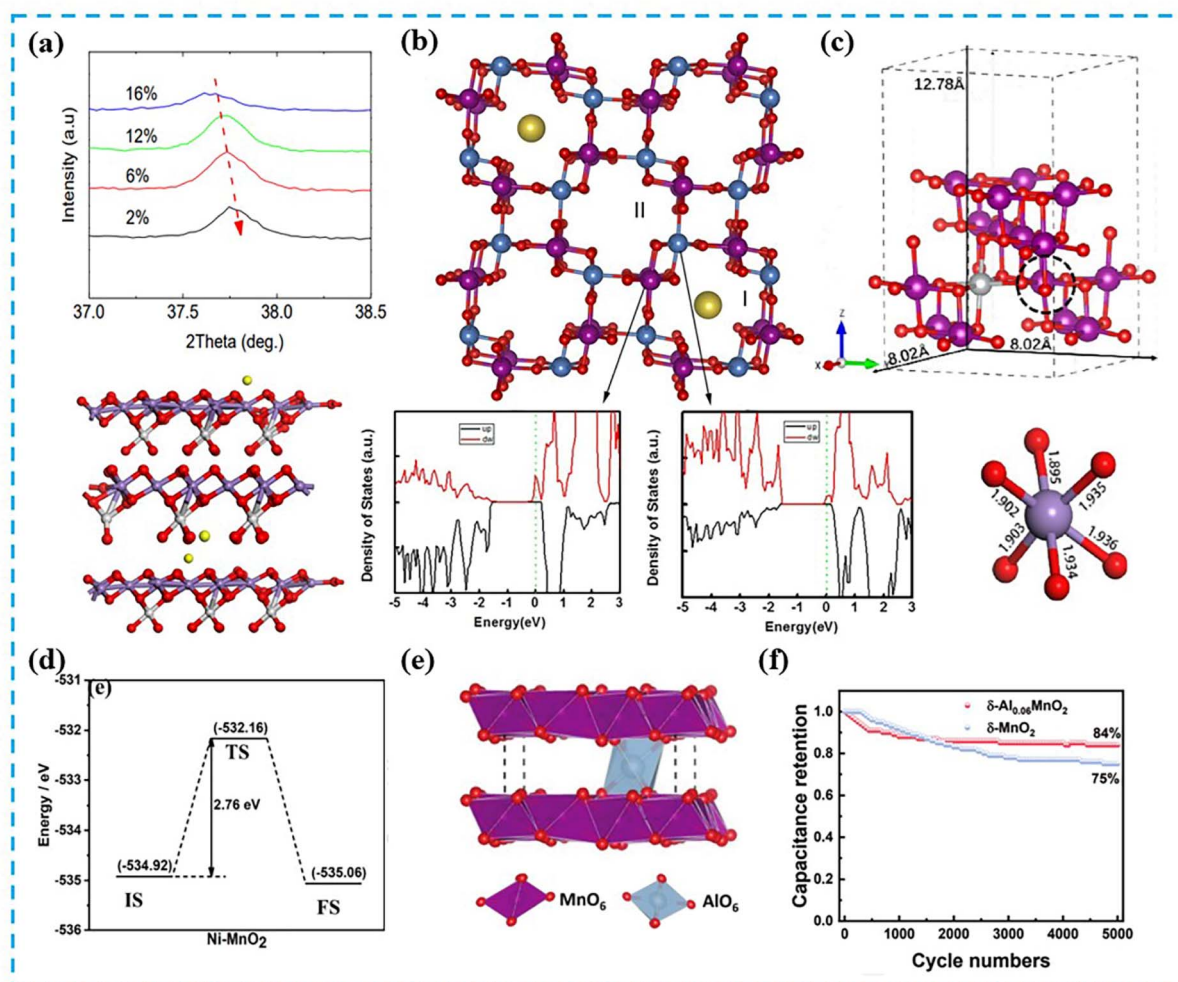


Fig. 9 (a) The magnification of the (211) crystal plane in the XRD spectra of K^+ -doped MnO_2 at different doping concentrations. (b) K^+ -doped two different Mn sites combined with the vertex and plane positions of the octahedral O in the MnO_2 crystal structure (blue and purple) and effects on PDOS. Reproduced with permission from ref. 96. Copyright 2015, Springer Nature. (c) The theoretically optimized structural model of $Ni-MnO_2$ and the magnified Mn-O bond length from the unit cell in the black circle. The purple, red and silver spheres represent the Mn, O and Ni atoms, respectively. (d) In $Ni-MnO_2$ illustration of Na^+ migration pathways and corresponding potential barriers. Reproduced with permission from ref. 97. Copyright 2022, Elsevier B.V. (e) The most stable structure of $\delta-Al_{0.06}MnO_2$. (f) Capacitance retention of $\delta-MnO_2$ and $\delta-Al_{0.06}MnO_2$ when the current density is 10 mA cm^{-2} . Reproduced with permission from ref. 67. Copyright 2022, Royal Society of Chemistry.



example, Wan *et al.*⁶⁷ doped Al into δ -MnO₂ and obtained high specific capacitance and cyclic stability. First-principles calculations showed that the Al-doped δ -MnO₂ was mainly interstitial doping to form a stable [AlO₆] octahedral structure (Fig. 9e). After Al doping, both the Mn 3d and O 2p orbitals in the valence band moved to the Fermi level, which led to the enhanced conductivity and provided more active centers for Li⁺ adsorption. In addition, the substitution of Al for Mn or O atoms could suppress the surface energy and hinder side reactions or structural transitions, thus achieving better cyclic stability (Fig. 9f). Poudel *et al.*²⁰ synthesized Gd-doped MnO₂ by hydrothermal method. The interlayer space was widened by the doping of rare earth metal cations with large ionic radius to form a unique structure, which improved the pseudocapacitive performance of the strategy. Furthermore, the doping of a large amount of Gd could also cause defects in the crystal, thus leading to increase in surface areas and the enhancement of kinetic reversibility.

Our research group also modified MnO₂-based materials by cationic doping. We have studied a variety of cationic doping, such as Ni, Co, Ce, Bi, *etc.* Among them, Ni, Co doping and Ce, Bi doping belong to substitution doping and interstitial doping, respectively. It was found that appropriate doping was beneficial to the improvement of the electrochemical performance of MnO₂, which was consistent with the results in the literature. However, not only the performance of MnO₂ could not be improved, but also the overall test results would be affected when the doping amount was slightly larger.

5.2.4 Anionic doping. Non-metal elements, such as S, N, B, P, *etc.*, have been shown to help narrow the wide band gap of MnO₂-based materials to improve their electrochemical performance.⁹⁸ In addition, it can help the material to build a porous framework because of its low negative charge, which can be used to obtain large specific surface areas and more surface active centers for promoting contacts between active materials and the electrolytes. For example, Chi and co-workers⁹⁹ prepared B-doped MnO₂ with 3D structure by chemical reaction. According to electrochemical test analysis, B-doped electrode material had higher specific capacitance or higher active material. Due to the electron-deficient nature of B³⁺, the concentration of OH⁻ on the surface of MnO₂ increased after B doping. The MnO₂ seeds formed by the disproportionation reaction between Mn⁷⁺ and Mn³⁺ were captured by OH⁻. And when B atoms entered the lattice, the difference of B, Mn and O ion radii could easily lead to lattice distortion and dislocation. So that lattice defects reduced the interfacial energy to speed up the rate capability. Up to now, there are few defect engineering strategies for anionic doping as MnO₂-based materials in SCs. There is no sufficient mechanistic explanation, and the lack of this aspect can provide more opportunities for its further exploration.

After experiments and research, it was found that defect engineering can change the internal structure of the MnO₂ lattice to a certain extent and improve its electronic structure. Using this strategy can achieve the improvement of electronic conductivity effectively, so as to achieve the purpose of improving the electrochemical performance of MnO₂.

5.3 Heterojunction engineering

Currently, the construction of heterojunctions is one of the most common methods for developing electrode materials. Manipulate electronic structure, enhance electronic/ionic conductivity, facilitate charge transfer kinetics, build porous structures, and improve structural stability by combining with other species.^{100–102} So, heterojunction engineering is also one of the most effective strategies to improve the adsorption and electrochemical performance of MnO₂-based materials. Recently, many MnO₂-based heterojunction materials for energy storage have been extensively reported.

5.3.1 MnO₂/metals. The heterojunction material has high specific capacitance and energy density as well as excellent cycling stability which formed by metal and MnO₂. It is one of the methods to obtain the best electrode materials for high-performance SCs. Surface modification/functionalization of metal nanoparticles for MnO₂-based materials can enhance the charge storage capacity. And noble metals (such as Au, Ag, Pt) can provide higher conductivity than ordinary metals.

For example, Luo *et al.*¹⁰³ utilized magnetron sputtering to synthesis the 3D network of nanofibres coated with gold on the surface. It could provide a large specific surface area and high porosity to improve the overall electrochemical performance of the material when MnO₂ was electrodeposition on the conductive networks. In another example, Khandare *et al.*¹⁰⁴ used gold nanoparticles to modify the surface of MnO₂ nanowires (Fig. 10a). Through surface functionalization and the increase of defects, the gold nanoparticles which decorated on MnO₂ nanowires provided good contacts for MnO₂, so that led to larger surface area and better storage capacity. The obtained cyclic stability of MnO₂ nanowires decorated with gold nanoparticles were better than MnO₂ nanowires (Fig. 10b). Deng *et al.*⁴⁷ constructed a 3D macroporous Mn core-shell structure through anodization process (Fig. 10c). Due to the high conductivity of the Mn layer, it was able to provide electron channels for charge storage and discharge (Fig. 10d). Moreover, Kumar *et al.*⁴⁴ reported the synthesis of nanoporous MnO₂-Cu structures by a facile method utilizing surface deformation and selective dealloying in the Cu-Mn system (Fig. 10e). The severe surface deformation before selective dealloying led to the *in situ* growth of nanoporous MnO₂ in the Cu-rich matrix, which enhanced the reaction kinetics. Furthermore, the nanoporous MnO₂-Cu structure provided a unique microstructure to ensure high specific surface area and small charge transfer resistance.

5.3.2 MnO₂/metal oxides. In recent years, transition metal oxides have received extensive attention because of their low price and easy availability, and they have become candidate materials for the study of electrode materials. MnO₂/metal oxide semiconductor heterojunctions, for instance, MnO₂/Mn₂O₃, MnO₂/Mn₃O₄, MnO₂/CuO, MnO₂/ZnO, MnO₂/TiO₂, MnO₂/Co₃O₄, MnO₂/MoO₃, *etc.*, have been successfully used to prepare electrode materials in SCs. Lu and co-workers¹⁰⁵ constructed hierarchical Mn₂O₃/MnO₂ core-shell nanofibers based on electrospinning technology (Fig. 11a). Among them, the Mn₂O₃ nanofibers provided a convenient transport path for the charge and a large number of sites for the MnO₂ core. The



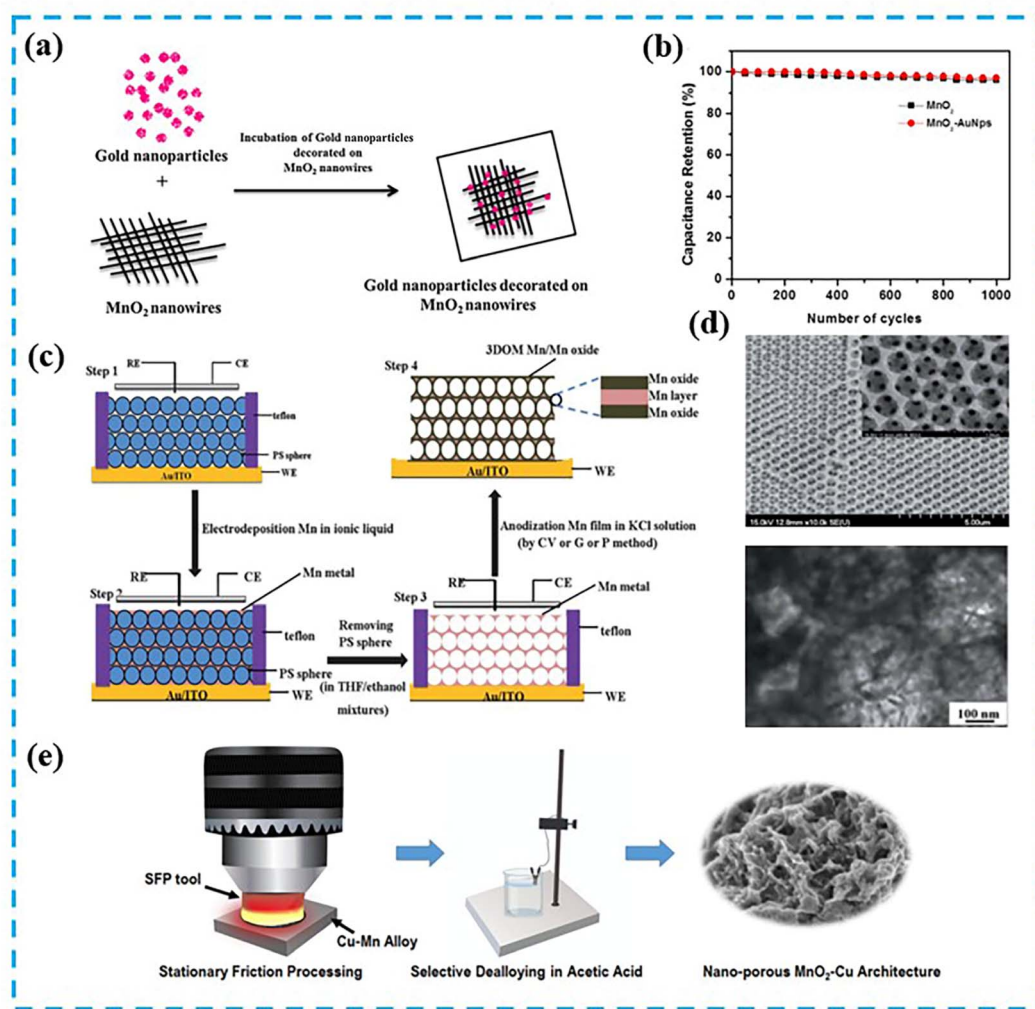


Fig. 10 (a) Illustration showing the steps in the process of MnO₂ nanowires decorated with gold nanoparticles. (b) Cyclic stability of MnO₂ nanowires decorated with gold nanoparticles. Reproduced with permission from ref. 104. Copyright 2017, Elsevier B.V. (c) Fabrication of highly porous 3D ordered macroporous Mn/Mn oxide electrodes. (d) SEM image (inset: partially enlarged) and TEM image of the inverse opal Mn with structure of 3D macro-porous. Reproduced with permission from ref. 47. Copyright 2013, Royal Society of Chemistry. (e) Development process of nanoporous MnO₂-Cu structures. The process included alloy synthesis by using an electric arc furnace, and severe surface deformation by using fixed friction machining and selective dealloying. Reproduced with permission from ref. 44. Copyright 2021, Elsevier Ltd.

formation of MnO₂ nanosheets reduced the length of the ion diffusion path and ensured the highest utilization of the material. Meanwhile, the core-shell heterogeneous structure further enhanced the Faraday redox reaction kinetics of the electrode material by increasing the specific surface area and improved the capacitive charge storage. Kang *et al.*¹⁰⁶ induced MnO₂ nanowires exfoliation to synthesize MnO₂/Mn₃O₄ hetero-junction materials. The electrochemical performance of a single MnO₂-based electrode material was improved due to the interaction between two active materials with different structures by combining the excellent electrical conductivity of 1D α -MnO₂ nanowires with the large surface area and abundant pseudocapacitive sites provided from external Mn₃O₄. Huang *et al.*¹⁰⁷ combined Kirkendall growth and Ostwald ripening processes to construct a CuO/MnO₂ core-shell structure without using any surfactant (Fig. 11b). CuO nanotubes were

formed by Kirkendall-type diffusion process. The interconnected MnO₂ nanosheets and hollow CuO nanotubes formed a highly porous state, which provided high surface areas and a large number of active sites for Na⁺ adsorption. The unique structure consisting of MnO₂ nanosheets coated with hollow CuO nanotubes also promoted ion transport and mechanical stability. Thus, the final electrochemical properties of the material were boosted. Huang *et al.*¹⁰⁸ studied the structure of the hierarchical ZnO/MnO₂ core-shell as well. The combination of ZnO and MnO₂ took advantage of the small size of ZnO and its ability to act as an effective electron transport pathway. The ZnO column arrays served as conductive scaffolds so that the subsequently grown MnO₂ nanosheets could avoid conventional aggregation and ensure sufficient charge transfer. The formed MnO₂ porous ultrathin nanosheets facilitated ion diffusion and improved energy storage capacity. Rajagopal



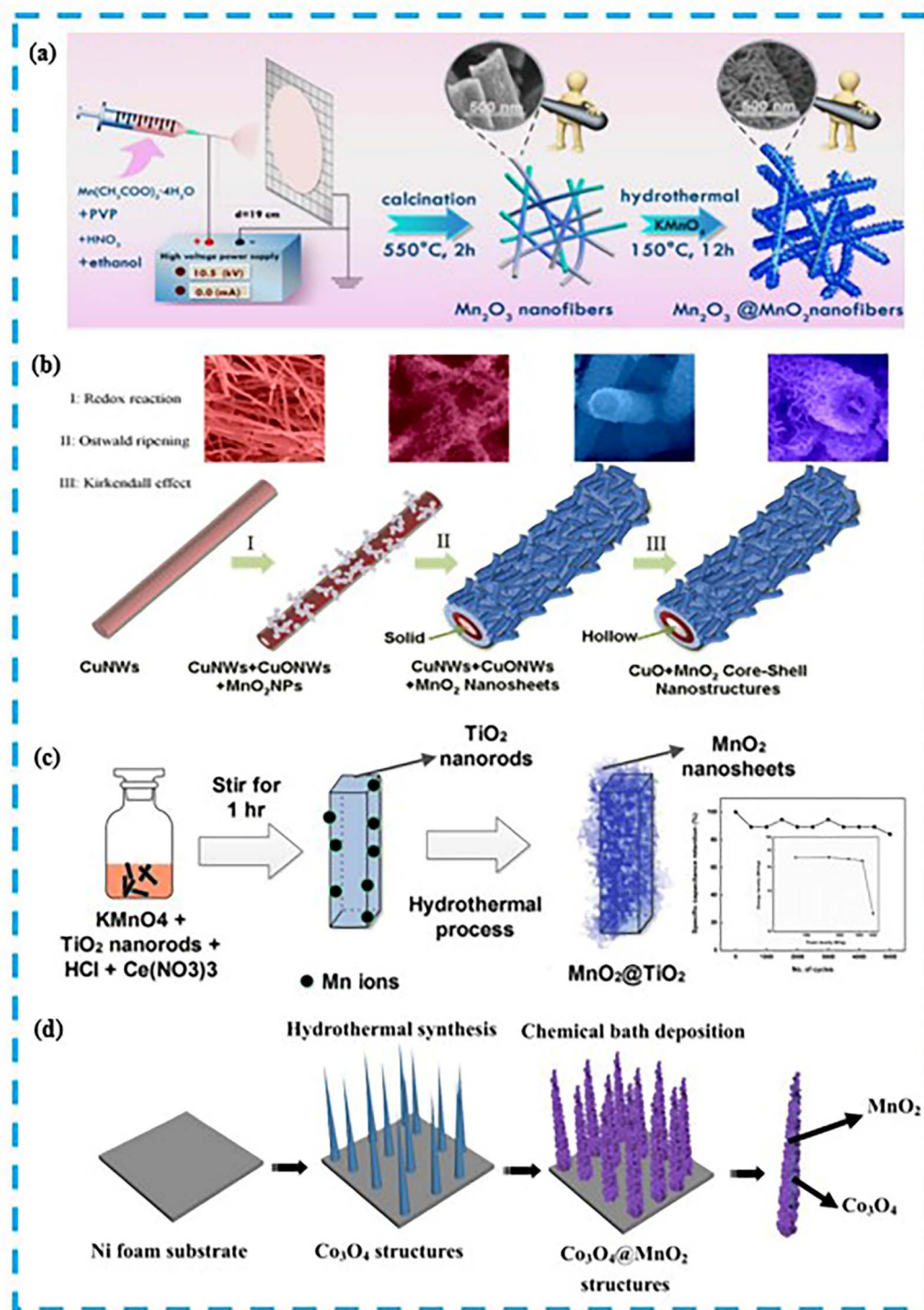


Fig. 11 (a) Diagram of the formation of $\text{Mn}_2\text{O}_3/\text{MnO}_2$ nanofibers. Reproduced with permission from ref. 105. Copyright 2020, American Chemical Society. (b) Schematic diagram of the growth mechanism of CuO/MnO_2 core-shell structure. Reproduced with permission from ref. 107. Copyright 2014, Springer Nature. (c) Schematic of formation $\text{TiO}_2/\text{MnO}_2$ core-shell structure, the inset shows the cycling stability of $\text{TiO}_2/\text{MnO}_2$ constructed ASC. Reproduced with permission from ref. 109. Copyright 2020, Elsevier B.V. (d) Diagram representation of the growth of 3D $\text{Co}_3\text{O}_4/\text{MnO}_2$ heterostructures grown on nickel foam. Reproduced with permission from ref. 111. Copyright 2016, Royal Society of Chemistry.

*et al.*¹⁰⁹ fabricated MnO_2 nanosheets packed with TiO_2 nanorods without using any template (Fig. 11c). The core-shell structured metal oxide nanocomposites provided high specific surface

areas as well as excellent porosity, which was beneficial for the enhancement of electrochemical energy storage. The constructed asymmetric supercapacitors (ASCs) also showed good



cycling stability. Nevertheless, Cao *et al.*¹¹⁰ changed the carrier concentration in TiO₂ by hydrogenation to improve its conductivity. Similar to the above-mentioned structural composition, the synthesized ultrathin MnO₂ coatings were covered on hydrogenated TiO₂ as the core to acquire MnO₂/H-TiO₂ materials with high capacitance and high cycling stability. Zheng *et al.*¹¹¹ prepared 3D Co₃O₄/MnO₂ heterostructures on flexible substrates (Fig. 11d). The layered structure of MnO₂ nanosheets provided large surface areas for fast ion diffusion and close electrode/electrolyte contact. This not only increased the effective pathway for ion and electron transport, but also strengthened the electrochemical behavior of the material. Furthermore, similar structural and electrochemical results could be obtained by using other methods.¹¹² Shafi *et al.*¹¹³ synthesized α -MnO₂/h-MoO₃ composites for the first time. The h-MoO₃ nanocrystals were presented as rods which surrounded by MnO₂ fiber structures consisting of needles. The ionic and electronic conductivity of the composites was enhanced by the heterogeneous structure formed by the tight bonding between the MnO₂ and MoO₃ nanorods. While MnO₂ provided large surface areas for adsorption and intercalation of charged ions to

greatly enhance the electrochemical performance of the material.

5.3.3 MnO₂/carbon materials. MnO₂-based nanomaterials have interesting structural and transport properties, but they have a serious barrier in electrode applications because of their poor electrical conductivity formed by their own electronic structure. Carbon materials are able to provide high power density and electrochemical stability. Composites with heterogeneous structures which prepared from carbon materials with MnO₂-based nanomaterials can overcome the drawback of poor electrical conductivity of MnO₂. Graphene as one of the 2D nanostructures of carbon materials is a monolayer material consisting of carbon atoms wrapped in a two-dimensional honeycomb sp² carbon lattice.¹¹⁴ It has many advantages, such as larger specific surface area and volume ratio, excellent electron transport properties, and chemical stability.¹¹⁵ Jangu *et al.*¹¹⁶ synthesized a unique V-shaped MnO₂/rGO material by a microwave-assisted hydrothermal method in a mildly acidic environment. It was found that the acetic acid concentration played an important role in the control of the V-shaped MnO₂ nanorod morphology (Fig. 12a). The V-shaped nanostructures

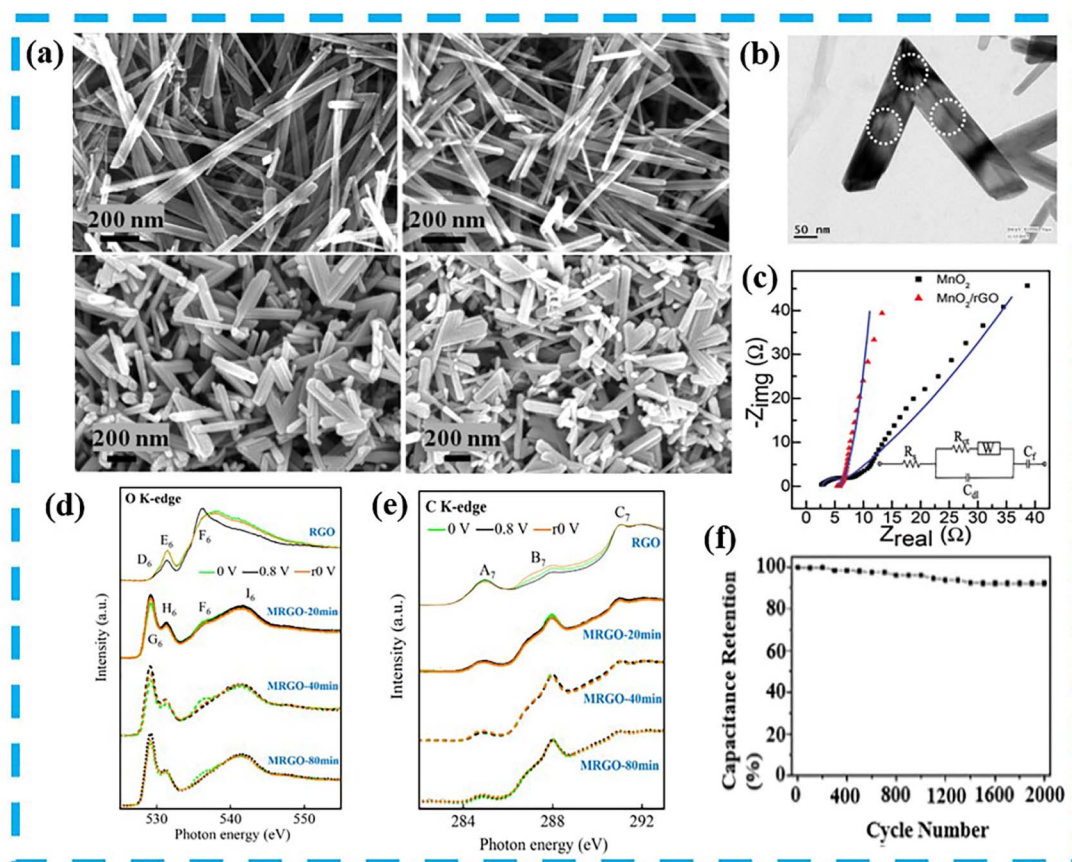


Fig. 12 (a) FESEM images of α -MnO₂ synthesized with different concentrations of acetic acid. (b) SEM image of V-shaped MnO₂ nanorods. (c) Energy quist plots for the comparison of MnO₂ and MnO₂/rGO in the frequency range of 0.01 Hz to 0.1 MHz. Reproduced with permission from ref. 116. Copyright 2021, Royal Society of Chemistry. (d) *Ex situ* XAS at the O K-edge of RGO and MRGO at different times under sequential application of potentials of 0, 0.8 and r0 V. Reproduced with permission from ref. 117. Copyright 2022, Elsevier B.V. (e) *Ex situ* XAS at C K-edges of RGO and MRGO at different times under sequential application of potentials of 0, 0.8 and r0 V. Reproduced with permission from ref. 117. Copyright 2022, Elsevier B.V. (f) Cyclic stability of the electrode at a current density of 0.2 A g⁻¹ for GCD. Reproduced with permission from ref. 118. Copyright 2020, AIP Publishing.



provided more space, which is beneficial for the electrolyte ions to diffuse (Fig. 12b). The existence of rGO reduced the charge transfer resistance in Fig. 12c. Chang *et al.*¹¹⁷ conducted a more in-depth study on the capacitive properties and electronic structure of MnO₂/rGO (MRGO) materials by means of electrochemistry and X-ray absorption spectroscopy (XAS). Electrochemical results indicated that the specific capacitances of RGO and MRGO were the result of the combined action of EDLC and

pseudo-capacitance. The XAS results demonstrated that the contribution of the RGO contrast capacitance involved the pseudo-capacitance and the related contribution of the O function in the EDLC (Fig. 12d and e). However, compared with pristine graphene, its electrical conductivity and cycling stability were lower due to the abundant topological defects of hydrophobic rGO sheets.⁴⁹ Thus, Bai *et al.*¹¹⁸ synthesized 3D graphene/MnO₂ composites *in situ* by a combination of

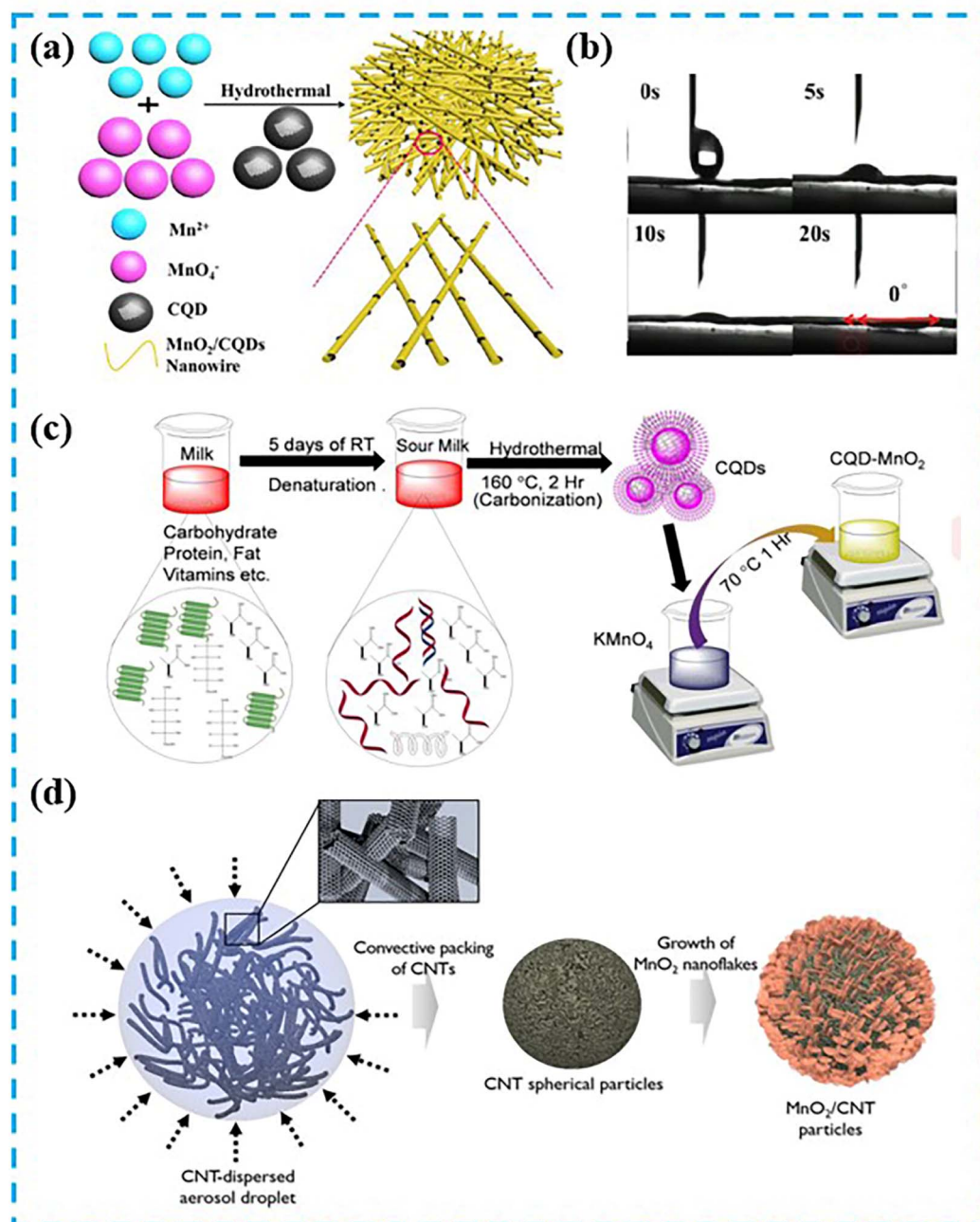


Fig. 13 (a) Schematic diagram of the fabrication process of CQD-induced MnO₂ nanowires composite films. (b) Images of CQD/MnO₂ composite film and the shape of water droplets on the film at 0, 5, 10 and 20 s, respectively. Reproduced with permission from ref. 120. Copyright 2017, American Chemical Society. (c) Schematic representation of the synthesis of CQD and CQD/MnO₂ nano hybrid materials. Reproduced with permission from ref. 121. Copyright 2018, Wiley-VCH. (d) Preparation diagram of CNT/MnO₂ core-shell particles. Reproduced with permission from ref. 123. Copyright 2017, American Chemical Society.



chemical vapor deposition and hydrothermal methods. The porous structure of 3D conducting graphene foam provided space for the rapid movement of electrolyte ions on the MnO_2 surface. The MnO_2 nanoflowers not only effectively restrained the collapse of pores in 3D graphene foam, but also formed layered structures that shortened the ion diffusion paths as well as enhancing the cyclic stability of the material (Fig. 12f).

Carbon Quantum Dots (CQDs) are considered as promising active nanomaterials due to their unique specific surface area, good water dispersibility and electrical conductivity.¹¹⁹ CQDs exhibit excellent electrochemical performance when combined with pseudocapacitive materials. Lv *et al.*¹²⁰ fabricated ultralong MnO_2 nanowires *via* a CQD-induced process (Fig. 13a). The formed CQD/ MnO_2 films exhibited superhydrophilicity in aqueous solution, which significantly improved the wettability between electrode and electrolyte (Fig. 13b). The obtained electrode material showed high specific capacitance, good reversibility and good stability. Prasath *et al.*¹²¹ prepared CQDs-coated MnO_2 nanostructures by using waste yogurt (Fig. 13c). It had excellent electrochemical performance due to its excellent electrical conductivity, extremely small size and carrying a variety of functional groups on its surface. In addition, carbon nanotubes (CNTs) also had higher electrical conductivity and controllable specific surface areas. Li *et al.*¹²² prepared 3D CNTs by large-area ultrasonic spraying of the surface, and the 3D CNTs provided sufficient sites for MnO_2 nanoparticles. High specific capacitance and good cycling stability were obtained through the good dispersion of MnO_2 on the CNTs surface. Gueon and co-workers¹²³ synthesized core-shell structures of CNTs coated with MnO_2 nanosheets (Fig. 13d). Because of the stronger binding of MnO_2 nanosheets on the CNT surface, this led to their low interfacial resistance. And the radial growth of MnO_2 nanosheets made the existence of radial spaces between the layers on the spherical CNTs particles, which enhanced the ion diffusion to obtain high capacitive performance (370 F g^{-1} at 0.5 A g^{-1}). Activated carbon (AC) is also commonly used for

hybridization with transition metal oxides. Choi *et al.*²⁴ used a simple hydrothermal method to synthesize AC/ MnO_2 composites with different mass ratios. The study found that the addition of MnO_2 could improve the electrochemical performance of AC within a certain range. However, with the increase of MnO_2 addition led to the decrease of porosity, which further made the electrochemical performance of the composites decreased.

5.3.4 MnO_2 /others. We have summarized other MnO_2 -based materials containing heterojunctions, as shown in Fig. 14. It can also be used to prepare interfacial structures, modify electronic structures, functionalize active sites and provide additional functionality.

Zhao *et al.*¹²⁴ prepared MnO_2 /carbon nano-support sheet (CNS) arrays *via* Mn-MOF-derived CNS nanosheets (Fig. 15a). The pristine framework of Mn-MOF not only provided a path for fast electron transport and enhanced electrical conductivity. The *in situ* formed MnO_2 nanosheet arrays and the derived carbon nanosheets also generated strong coupling, so the final electrode material achieved excellent capacitance and cyclic stability. Chen *et al.*¹²⁵ *via* Fe_2O_3 derived from MIL-88A as a core and surface-modified it with an array of vertically grown MnO_2 nanosheets. The as-synthesized Fe_2O_3 @ MnO_2 material provided a sufficient interface for electrode-electrolyte contact due to the modification of nanosheet arrays. The large pore volume also provided portability for rapid ion diffusion, so it exhibited good capacitive performance (908.5 F g^{-1} at 1 A g^{-1}) and cyclic stability. Fu *et al.*¹²⁶ prepared core/sheath structured MnO_x /PPy nanowires *via in situ* polymerization of polypyrrole (PPy) (Fig. 15b). Conductive V_O was introduced into MnO_x simultaneously during the formation of the PPy sheath. And the mesoporous PPy nanoparticles were used to form chemical bonds with their underlying layers to promote electron transfer and restraint the dissolution of Mn. Liu *et al.*¹²⁷ obtained MnO_2 nanosheet structures enriched with 3D mesopores interconnected with polyaniline (PANI) chains, which underwent

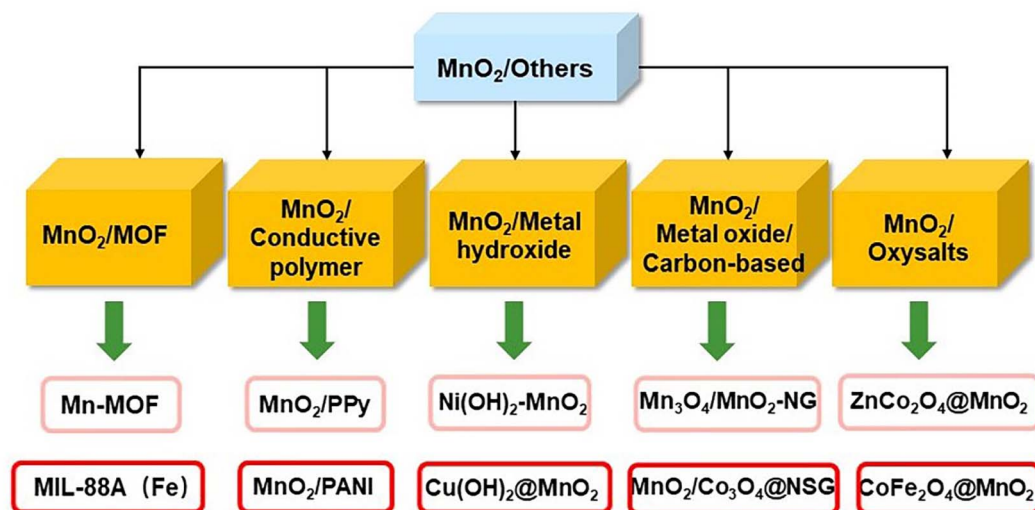


Fig. 14 Recently reported heterogeneous junctions of MnO_2 -based electrode materials with other substances.



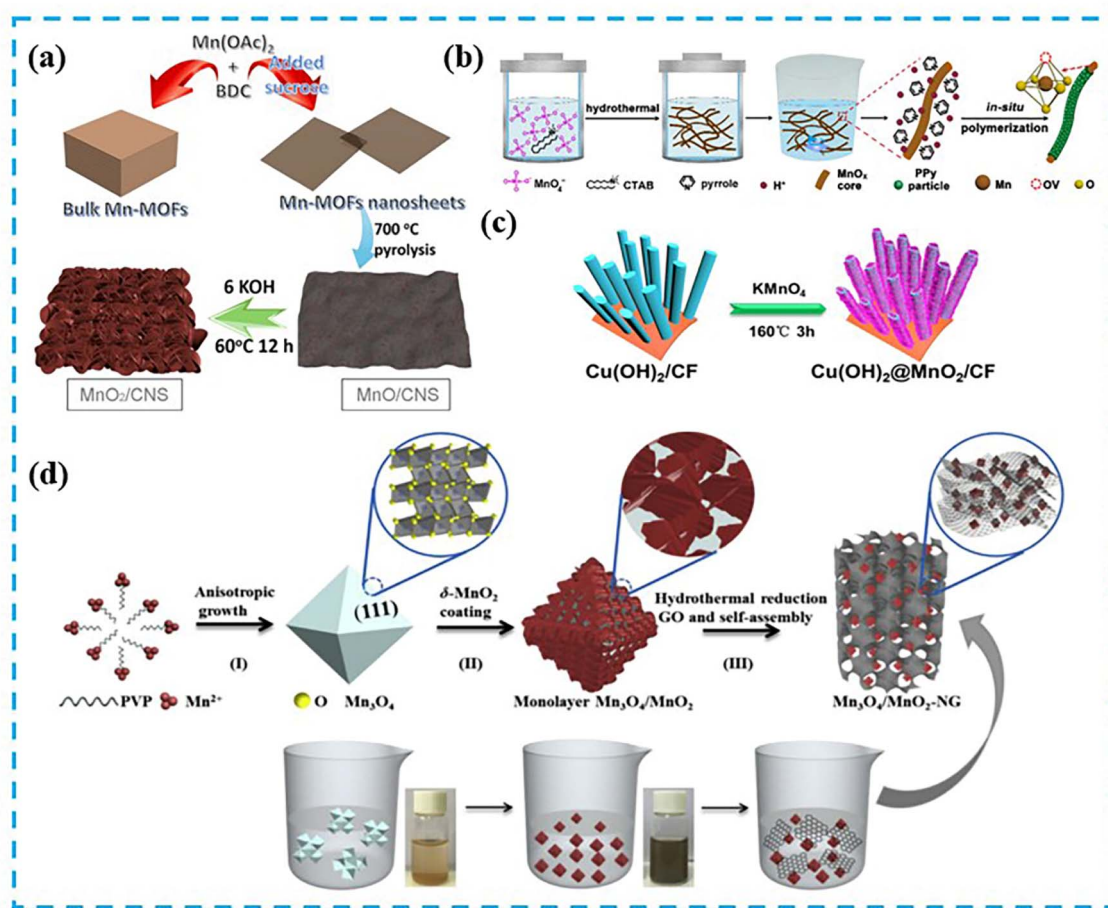


Fig. 15 (a) Schematic diagram of the preparation of vertically aligned MnO₂ nanosheets strongly coupled with carbon nanosheets. Reproduced with permission from ref. 124. Copyright 2018, Springer Nature. (b) Illustrative diagram of the formation of MnO₂/PPy nanowires. Reproduced with permission from ref. 126. Copyright 2019, Elsevier Inc. (c) Schematic diagram of the synthesis of Cu(OH)₂@MnO₂/CF nanorod arrays. Reproduced with permission from ref. 129. Copyright 2019, Elsevier Inc. (d) Synthesis in process of Mn₃O₄/MnO₂ octahedral, as well as Mn₃O₄/MnO₂ particles and Mn₃O₄/MnO₂-NG. Reproduced with permission from ref. 130. Copyright 2018, Elsevier B.V.

a fast chemical oxidation polymerization reaction by electrostatically grafting the negatively charged surface of MnO₂ with aniline monomers. Due to the exposure of the reaction center on the electrode surface and the porous structure formed by the combination of PANI chains, it could be used for electron transport and ion diffusion pathway shortening. Li *et al.*¹²⁸ synthesized Ni(OH)₂-MnO₂ hierarchical nanosheet arrays. The formed heterostructure provided lattice fringe distortion as additional active centers, which combined the synergistic effect of the two to improve the capacitive performance (area capacitance of 14.7 F cm⁻² at 2.5 mA cm⁻²) of the material. Wang *et al.*¹²⁹ synthesized Cu(OH)₂@MnO₂ core-shell nanorod arrays on a copper foam substrate (Fig. 15c). The 1D Cu(OH)₂ nanocrystalline nuclei provided a scaffold for MnO₂ nanosheet growth and facilitated the ions and electrons to diffuse rapidly. The 2D MnO₂ nanosheets used as the shell provided a great deal of active sites. Cui *et al.*¹³⁰ reported novel Mn₃O₄/MnO₂-NG hierarchical nanostructures (Fig. 15d). The 3D conductive network of NG provided an open path for electron/ion transport while being in close contact with Mn₃O₄/MnO₂. The enhanced

conductivity also increased the synergy of the components, which promoted the electrode reaction kinetics. Adaikalam *et al.*¹³¹ improved the problems of poor reversibility and stability in the presence of single MnO₂ and Co₃O₄ electrodes by N and S double doped GO sheets. The overall mechanical strength and electrical conductivity of the material was also enhanced and allowed more ion penetration into the electrode. Hybridization of MnO₂-based materials with oxygenates could enhance the electrochemical performance through their structural advantages. Gao *et al.*¹³² designed unique CoFe₂O₄@MnO₂ nanoarrays based on nickel foam. The porous framework formed by CoFe₂O₄ contributed sites for the growth of MnO₂, which promoted the enhancement of electronic/ionic conductivity. Jia *et al.*¹³³ prepared ZnCo₂O₄@MnO₂ core-shell structure in which the generation of MnO₂ films could provide large specific surface areas to efficiently promote the rapid ion transport. The final result is reflected in the assembled ASC devices. Therefore, it shows different MnO₂-based materials glowing LEDs for heterojunction engineering in Fig. 16.





Fig. 16 (a) Device of Ar-plasma treated α -MnO₂ nanowires in two series powered 65 LEDs. Reproduced with permission from ref. 134. Copyright 2018, Elsevier B.V. (b) The V-shaped MnO₂ powers four LEDs after charging. Reproduced with permission from ref. 116. Copyright 2021, Royal Society of Chemistry. (c) Digital image of a red LED lit by the ASC composed of MnO₂/CNT. Reproduced with permission from ref. 123. Copyright 2017, American Chemical Society. (d) A red and a yellow LED illuminated by the single device of hierarchical MnO₂-Mn₃O₄. Reproduced with permission from ref. 106. Copyright 2019, Elsevier Ltd. (e) Photographic images of two solid-state ASCs made of Cu(OH)₂@MnO₂/CF in series lighting a red LED bulb. Reproduced with permission from ref. 129. Copyright 2019, Elsevier Inc. (f) A digital image showing the two devices composed of CoFe₂O₄@MnO₂ in series can lighten up ten red LED indicators. Reproduced with permission from ref. 132. Copyright 2017, Elsevier Ltd.

The systematically summarizes the adjustment of the electrochemical properties of MnO₂-based electrode materials by different engineering strategies, and shows the differences of each strategy in Table 2. Overall, the results obtained are still far from the theoretical requirements.

6. Summary and perspectives

MnO₂ is a potential electrode material in both crystal structure control and energy storage. In this review, we give a systematic description of various strategies based on MnO₂-based electrode materials through morphology engineering, defect engineering and heterojunction engineering. Many strategies are used to prepare MnO₂-based electrode materials to improve their electrochemical properties. However, there are still problems that deserve our attention, whether using strategies such as morphology engineering or defect engineering, which are only matter of fact, all lacking a systematic process and mechanism analysis of a single strategy. As a conclusion, MnO₂-based electrode materials still have numerous difficulties to overcome, while leaving some space for more in-depth exploration (Fig. 17).

The following are issues that may need to be solved immediately:

(1) The morphology engineering based on MnO₂ nanostructures can be used to influence the final electrochemical performance by means of extrinsic modification. For example, increasing specific surface area and exposing effective active sites, shortening ion diffusion paths and electron transport distances, and promoting charge transfer and accelerating

reaction kinetics.^{81–83} At present, a large number of synthetic methods on the morphology and porous structure of MnO₂ have been developed to enhance its electrochemical behavior. Most studies are limited to the same dimension, and they lack systematic studies on the formation of MnO₂ porous structure in multi-dimensional space. In the process of electrochemical reaction, the existence of disproportionation reaction will lead to the reduction of specific surface area or the reduction of effective active sites in the MnO₂ structure. Moreover, the more complex structure of MnO₂ also requires a complete set of theories to describe the surface reaction mechanisms corresponding to different morphologies. Therefore, this type of research is relatively lacking, and new methods need to be explored to obtain structures with excellent properties and certain stability. We should focus on the reports which related to new technologies and methods to stabilize the morphology and structure of MnO₂. In addition, it is quite difficult to obtain high-purity crystalline MnO₂, and it is urgent to develop new strategies to prepare single-phase MnO₂ materials.

(2) Defect engineering can adjust the electronic structure of MnO₂-based materials effectively according to the formation mechanisms of different defects. For instance, oxygen vacancy can change the bond length by adjusting the state of the active centers. And cation vacancies can provide additional intercalation sites to enhance the storage capacitance. Besides, cationic doping and anionic doping can use positive and negative charge attraction and different positions in the lattice to change the structure band gap, thereby enhancing conductivity and releasing electrochemical activity.^{88,89,93,98} There are few reports on anionic doping as MnO₂-based electrode materials, so that



Table 2 Differences in electrochemical performance of MnO₂-based electrode materials displayed by different engineering strategies

Strategy	Active materials	Cycle stability	Specific capacitance	Electrolyte	Rate stability	Ref.
Morphology engineering	MnO ₂ nanospheres thin films	72.3% (after 1000 cycles)	206.2 F g ⁻¹ at 1.25 A g ⁻¹	1 M Na ₂ SO ₄	75.8% (1.25 to 5 A g ⁻¹)	71
	α-MnO ₂ nanorods	87% (after 100 cycles)	152 F g ⁻¹ at 2.5 mA	1 M Na ₂ SO ₄	75% (2.5 to 10 mA)	72
	α-MnO ₂ nanowires	78% (after 2000 cycles)	180 F g ⁻¹ at 1 A g ⁻¹	1 M Na ₂ SO ₄	40% (1 to 8 A g ⁻¹)	73
	α-MnO ₂ hollow spheres and hollow urchins	89% (after 350 cycles)	167 F g ⁻¹ at 2.5 mA	1 M Na ₂ SO ₄	74.3% (2.5 to 10 mA)	74
	δ-MnO ₂ /GP	87.2% (after 30 000 cycles)	446.6 F g ⁻¹ at 1 A g ⁻¹	1 M NaNO ₃	66.8% (1 to 10 A g ⁻¹)	75
Defect engineering	Defective MnO ₂	89.4% (after 4000 cycles)	202 F g ⁻¹ at 1 A g ⁻¹	1 M Na ₂ SO ₄	NA	79
	MnO ₂ -C	83.9% (after 10 000 cycles)	286.9 F g ⁻¹ at 1 A g ⁻¹	1 M Na ₂ SO ₄	78.4% (1 to 20 A g ⁻¹)	80
	2D birnessite δ-MnO ₂	84.7% (after 5000 cycles)	382 F g ⁻¹ at 5 mV s ⁻¹	1 M KOH	67.4% (1 to 8 A g ⁻¹)	81
	MnO _x @rGO microspheres	90% (after 10 000 cycles)	405 F g ⁻¹ at 1 A g ⁻¹	6 M KOH	76.5% (1 to 20 A g ⁻¹)	82
	δ-MnO ₂ nanosheets	83% (after 1000 cycles)	306 F g ⁻¹ at 0.2 A g ⁻¹	1 M Na ₂ SO ₄	NA	87
	Ni-MnO ₂	85% (after 10 000 cycles)	379 F g ⁻¹ at 1 A g ⁻¹	1 M Na ₂ SO ₄	74.1% (1 to 20 A g ⁻¹)	89
	Al doped δ-MnO ₂	84% (after 5000 cycles)	450 mF cm ⁻² at 0.1 mA cm ⁻²	1 M Li ₂ SO ₄	77.8% (0.1 to 2 mA cm ⁻²)	90
	GdMnO ₂ /Ni(OH) ₂ nanorods	108.3% (after 10 000 cycles)	118.61 mA h g ⁻¹ at 1 A g ⁻¹	2 M KOH	57.5% (1 to 10 A g ⁻¹)	20
	B doped MnO ₂ film	NA	269 F g ⁻¹ at 50 mV s ⁻¹	0.5 M Na ₂ SO ₄	58.7% (2 to 500 mV s ⁻¹)	92
Heterojunction engineering	Au-MnO ₂ nanowires	97% (after 1000 cycles)	249 F g ⁻¹ at 1 mV s ⁻¹	1 M Na ₂ SO ₄	65.9% (1 to 5 mV s ⁻¹)	96
	Nano-porous MnO ₂ -Cu architecture	95% (after 4000 cycles)	2.8 F cm ⁻² at 5 mA cm ⁻²	1 M Na ₂ SO ₄	35.7% (5 to 25 mA cm ⁻²)	97
	Mn ₂ O ₃ @MnO ₂ core-shell nanofibers	86.5% (after 5000 cycles)	225 F g ⁻¹ at 0.2 A g ⁻¹	1 M Na ₂ SO ₄	39.8% (0.1 to 10 A g ⁻¹)	98
	Hierarchical MnO ₂ -Mn ₃ O ₄	99.3% (after 5000 cycles)	406.2 F g ⁻¹ at 1 A g ⁻¹	1 M Na ₂ SO ₄	68.8% (1 to 20 A g ⁻¹)	99
	Nanoflakes MnO ₂ /rGO	99.4% (after 5000 cycles)	482.4 F g ⁻¹ at 10 mV s ⁻¹	1 M Na ₂ SO ₄	NA	100
	3D-graphene/MnO ₂ foam	92.2% (after 2000 cycles)	333.4 F g ⁻¹ at 0.2 A g ⁻¹	1 M Na ₂ SO ₄	NA	111
	CQD-MnO ₂	100% (after 1200 cycles)	189 F g ⁻¹ at 0.14 A g ⁻¹	1 M Na ₂ SO ₄	34.4% (0.14 to 2 A g ⁻¹)	114
	MnO ₂ /CNS nanosheet arrays	96.1% (after 5000 cycles)	339 F g ⁻¹ at 0.5 A g ⁻¹	1 M Na ₂ SO ₄	53.8% (0.5 to 2 A g ⁻¹)	117
	MnO _x /PPy nanowires core/sheath structure	97.4% (after 10 000 cycles)	1091.4 F g ⁻¹ at 1 A g ⁻¹	1 M Na ₂ SO ₄	81.6% (1 to 20 A g ⁻¹)	119
	Cu(OH) ₂ @MnO ₂ /CF core-shell nanorods array	85.17% (after 5000 cycles)	708.6 mF cm ⁻² at 2 mA cm ⁻²	6 M KOH	74.65% (2 to 20 mA cm ⁻²)	122
MnO ₂ /Co ₃ O ₄ with N and S co-doped GO	95% (after 10 000 cycles)	614 F g ⁻¹ at 1 A g ⁻¹	6 M KOH	57% (1 to 30 A g ⁻¹)	124	

the reaction mechanism of electrode materials in electrolyte is not perfect. Therefore, it is necessary to expand the research scope of anions to deeply understand and explore them in this regard.

(3) The construction of MnO₂-based heterojunction materials is one of the most effective strategies to enhance the electrochemical performance. The electronic structure is regulated by the interaction generated by binding with other substances, promoting the charge transfer kinetics, building



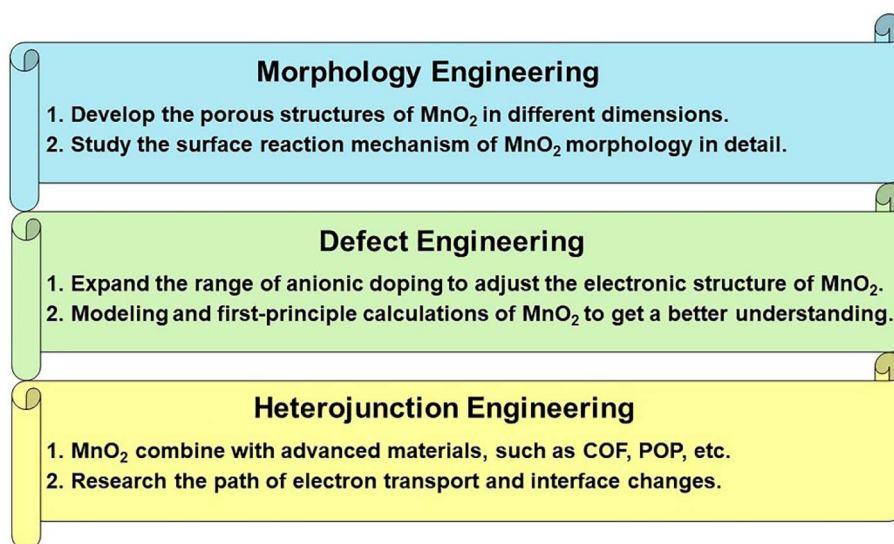


Fig. 17 Prospects for MnO₂-based electrode materials for SCs.

the porous structure and improving the stability of the structure.^{127–132} Up to now, great progress has been made and the exploration of heterojunctions is not over yet, so further efforts are still needed. In the future research process, it can focus on combining MnO₂ with advanced materials (such as COF, POP, etc.) to form heterojunction materials. Based on the above basis, advanced characterization methods are used to deeply explore the changes of its electron transport paths and interfaces.

In summary, the reaction mechanisms of MnO₂ as electrode materials for SCs need to be further explored and refined. We sincerely hope that the review will provide an overview of the study progress of MnO₂-based electrode materials for SC and theoretical guidance for subsequent basic and applied research.

Author contributions

Juyin Liu: project conceptualization, methodology, investigation, writing – original draft, writing – review & editing, visualization. Jiali Bao: writing – review & editing. Xin Zhang: writing – review & editing. Yanfang Gao: resources, funding acquisition. Yao Zhang: supervision. Ling Liu: writing – review & editing, resources, funding acquisition. Zhenzhu Cao: writing – review & editing.

Conflicts of interest

There are no conflicts to declare.

Acknowledgements

This work was supported by the National Natural Science Foundation of China (No. 21566030) and Ministry of Science and Technology China-South Africa Joint Research Program

(No. CS08-L15), Project of Inner Mongolia Education Department (NJZY089), Natural Science Foundation of Inner Mongolia (2015MS0205), Scientific Research Startup Foundation Project (No. DC2200000899), the National Natural Science Foundation of China (No. 22165021), Scientific Foundation of High Education Institutes Inner Mongolia of China (NJZY20071), and creative talents team of “Prairie Talent” engineering industry and the “Prairie Talent” of organization Department of Inner Mongolia Party Committee.

References

- 1 P. Simon and Y. Gogotsi, *Nat. Mater.*, 2008, 7, 845–854.
- 2 G. Ma, M. Dong, K. Sun, E. Feng, H. Peng and Z. Lei, *J. Mater. Chem. A*, 2015, 3, 4035–4041.
- 3 R. M. John and P. Simon, *Science*, 2008, 321, 651–652.
- 4 R. Wang, M. Yao and Z. Niu, Smart supercapacitors from materials to devices, *InfoMat*, 2020, 2, 113–125.
- 5 Y. Su, N. Li, L. Wang, R. Lin, Y. Zheng, G. Rong and M. Sawan, *Adv. Mater. Technol.*, 2022, 7, 2100608.
- 6 M. Shi, C. Yang, X. Song, L. Zhao, J. Liu, P. Zhang and L. Gao, *Adv. Sustainable Syst.*, 2017, 1, 1700044.
- 7 R. K. Sarojini, K. Palanisamy, P. Sanjeevikumar and J. B.-H. Nielsen, *IET Renew. Power Gener.*, 2020, 14, 1156–1163.
- 8 L. Chai, P. Wang, X. Liu, Y. Sun, X. Li and J. Pan, *J. Power Sources*, 2022, 532, 231324.
- 9 X. Gang, M. Krishnamoorthy, W. Jiang, J. Pan, Z. Pan and X. Liu, *Carbon*, 2021, 171, 62–71.
- 10 L. Sheng, L. Jiang, T. Wei, Z. Liu and Z. Fan, *Adv. Energy Mater.*, 2017, 7, 1700668.
- 11 H. Jin, X. Feng, J. Li, M. Li, Y. Xia, Y. Yuan, C. Yang, B. Dai, Z. Lin, J. Wang, J. Lu and S. Wang, *Angew. Chem., Int. Ed.*, 2019, 58, 2397–2401.



- 12 H. S. Kim, J. B. Cook, H. Lin, J. S. Ko, S. H. Tolbert, V. Ozolins and B. Dunn, *Nat. Mater.*, 2017, **16**, 454–460.
- 13 S. Liu, A. Li, C. Yang, F. Ouyang, J. Zhou and X. Liu, *Appl. Surf. Sci.*, 2022, **571**, 151306.
- 14 A. Muzaffar, M. B. Ahamed, K. Deshmukh and J. Thirumalai, *Renewable Sustainable Energy Rev.*, 2019, **101**, 123–145.
- 15 Z. K. He, Y. Lu, C. Zhao, J. Zhao, Z. Gao and Y. Y. Song, *Appl. Surf. Sci.*, 2021, **567**, 150832.
- 16 M. Y. Chung and C. T. Lo, *Electrochim. Acta*, 2020, **364**, 137324.
- 17 I. Ryu, D. Kim, G. Choe, S. Jin, D. Hong and S. Yim, *J. Mater. Chem. A*, 2021, **9**, 26172–26180.
- 18 K. C. Devarayapalli, K. Lee, H. B. Do, N. N. Dang, K. Yoo, J. Shim and S. V. P. Vattikuti, *Mater. Today Energy*, 2021, **21**, 100699.
- 19 Z. Shi, L. Xing, Y. Liu, Y. Gao and J. Liu, *Carbon*, 2018, **129**, 819–825.
- 20 M. B. Poudel and H. J. Kim, *J. Energy Chem.*, 2022, **64**, 475–484.
- 21 Z. Shi, J. Liu, Y. Gao and Y. Xu, *J. Mater. Sci.*, 2021, **56**, 1612–1629.
- 22 G. Manibalan, Y. Govindaraj, J. Yesuraj, P. Kuppusami, G. Murugadoss, R. Murugavel and M. R. Kumar, *J. Colloid Interface Sci.*, 2021, **585**, 505–518.
- 23 Y. Zhang, X. Yuan, W. Lu, Y. Yan, J. Zhu and T. W. Chou, *Chem. Eng. J.*, 2019, **368**, 525–532.
- 24 J. R. Choi, J. W. Lee, G. Yang, Y. J. Heo and S. J. Park, *Catalysts*, 2020, **10**, 256.
- 25 Y. Mao, J. Xie, C. Guo, H. Liu, H. Xiao and W. Hu, *Chem. Eng. J.*, 2021, **426**, 131188.
- 26 Y. Wei, M. Zheng, W. Luo, B. Dai, J. Ren, M. Ma, T. Li and Y. Ma, All pseudocapacitive MXene-MnO₂ flexible asymmetric supercapacitor, *J. Energy Storage*, 2022, **45**, 103715.
- 27 S. G. Mohamed, S. Y. Attia and H. H. Hassan, *Microporous Mesoporous Mater.*, 2017, **251**, 26–33.
- 28 N. R. Chodankar, D. P. Dubal, Y. Kwon and D.-H. Kim, *NPG Asia Mater.*, 2017, **9**, e419.
- 29 L. Lin, S. Tang, S. Zhao, X. Peng and N. Hu, *Electrochim. Acta*, 2017, **228**, 175–182.
- 30 H. Wu, Z. Lou, H. Yang and G. Shen, *Nanoscale*, 2015, **7**, 1921–1926.
- 31 M. Isacfranklin, S. Daphine, R. Yuvakkumar, L. Kungumadevi, G. Ravi, A. G. Al-Sehemi and D. Velauthapillai, *Ceram. Int.*, 2022, **48**, 24745–24750.
- 32 Q. Ma, F. Cui, J. Zhang and T. Cui, *J. Colloid Interface Sci.*, 2023, **629**, 649–659.
- 33 A. K. Das, N. H. Kim, S. H. Lee, Y. Sohn and J. H. Lee, *Composites, Part B*, 2018, **150**, 269–276.
- 34 A. K. Das, N. H. Kim, S. H. Lee, Y. Sohn and J. H. Lee, *Composites, Part B*, 2018, **150**, 234–241.
- 35 P. Zhang and H. He, *J. Alloys Compd.*, 2020, **826**, 153993.
- 36 Y. Dong, Y. Wang, Y. Xu, C. Chen, Y. Wang, L. Jiao and H. Yuan, *Electrochim. Acta*, 2017, **225**, 39–46.
- 37 X. Zheng, Y. Ye, Q. Yang, B. Geng and X. Zhang, *Dalton Trans.*, 2016, **45**, 572–578.
- 38 S. Nagamuthu, S. Vijayakumar, S.-H. Lee and K.-S. Ryu, *Appl. Surf. Sci.*, 2016, **390**, 202–208.
- 39 Q. Zhao, A. Song, S. Ding, R. Qin, Y. Cui, S. Li and F. Pan, Preintercalation Strategy in Manganese Oxides for Electrochemical Energy Storage: Review and Prospects, *Adv. Mater.*, 2020, **32**, 2002450.
- 40 J. Shin, J. K. Seo, R. Yaylian, A. Huang and Y. S. Meng, *Int. Mater. Rev.*, 2019, 1–32.
- 41 V. Veeramani, B. Dinesh, S. M. Chen and R. Saraswathi, *J. Mater. Chem. A*, 2016, **4**, 3304–3315.
- 42 S. B. Singh, T. I. Singh, N. H. Kim and J. H. Lee, *J. Mater. Chem. A*, 2019, **7**, 10672–10683.
- 43 S. Sun, G. Jiang, Y. Liu, B. Yu and U. Evariste, *J. Energy Storage*, 2018, **18**, 256–258.
- 44 A. Kumar, A. Thomas, A. Gupta, M. Garg, J. Singh, G. Perumal, E. Sujithkrishnan, P. Elumalai and H. S. Arora, *J. Energy Storage*, 2021, **42**, 103100.
- 45 K. Xiao, J. W. Li, G. F. Chen, Z. Q. Liu, N. Li and Y.-Z. Su, *Electrochim. Acta*, 2014, **149**, 341–348.
- 46 Y. Wang, Y. Wang and L. Jiang, *J. Appl. Electrochem.*, 2018, **48**, 495–507.
- 47 M. J. Deng, P. J. Ho, C. Z. Song, S. A. Chen, J. F. Lee, J. M. Chen and K. T. Lu, *Energy Environ. Sci.*, 2013, **6**, 2178–2185.
- 48 J. Cao, Y. Wang, Y. Zhou, J. H. Ouyang, D. Jia and L. Guo, *J. Electroanal. Chem.*, 2013, **689**, 201–206.
- 49 S. Sun, P. Wang, S. Wang, Q. Wu and S. Fang, *Mater. Lett.*, 2015, **145**, 141–144.
- 50 Z. Liu, J. Zhang, X. Yang, Y. He, Y. Bai, L. Kang, H. Xu, F. Shi and Z. Lei, *J. Mater. Chem. A*, 2016, **4**, 9088–9096.
- 51 C. Choi, H. J. Sim, G. M. Spinks, X. Lepró, R. H. Baughman and S. J. Kim, *Adv. Energy Mater.*, 2016, **6**, 1502119.
- 52 T. Park, Y. Jang, J. W. Park, H. Kim and S. J. Kim, *RSC Adv.*, 2020, **10**, 14007–14012.
- 53 M. Liu, L. Gan, W. Xiong, Z. Xu, D. Zhu and L. Chen, *J. Mater. Chem. A*, 2014, **2**, 2555–2562.
- 54 M. Li, J. Yu, X. Wang and Z. Yang, *Appl. Surf. Sci.*, 2020, **530**, 147230.
- 55 S. Devaraj and N. Munichandraiah, *J. Phys. Chem. C*, 2008, **112**, 4406–4417.
- 56 S. D. Huang, C. Shang, X. J. Zhang and Z. P. Liu, *Chem. Sci.*, 2017, **8**, 6327–6337.
- 57 W. Yao, G. M. Odegard, Z. Huang, Y. Yuan, H. Asayesh-Ardakani, S. Sharifi-Asl, M. Cheng, B. Song, R. Deivanayagam, F. Long, C. R. Friedrich, K. Amine, J. Lu and R. Shahbazian-Yassar, *Nano Energy*, 2018, **48**, 301–311.
- 58 T. Barudžija, V. Kusigerski, N. Cvjetičanin, S. Šorgić, M. Perović and M. Mitrić, *J. Alloys Compd.*, 2016, **665**, 261–270.
- 59 R. Yang, Y. Fan, R. Ye, Y. Tang, X. Cao, Z. Yin and Z. Zeng, *Adv. Mater.*, 2021, **33**, 2004862.
- 60 M. T. N. Dinh, C. C. Nguyen, T. L. T. Vu, V. T. Ho and Q. D. Truong, *Appl. Catal., A*, 2020, **595**, 117473.
- 61 X. Wang and Y. Li, *Chem. Commun.*, 2002, 764–765.
- 62 Y. Chabre and J. Pannetier, *Prog. Solid State Chem.*, 1995, **23**, 1–130.



- 63 J. Zhao, Z. Tao, J. Liang and J. Chen, *Cryst. Growth Des.*, 2008, **8**, 2799–2805.
- 64 S. L. Brock, N. Duan, Z. R. Tian, O. Giraldo, H. Zhou and S. L. Suib, *Chem. Mater.*, 1998, **10**, 2619–2628.
- 65 C. H. Kim, Z. Akase, L. Zhang, A. H. Heuer, A. E. Newman and P. J. Hughes, *J. Solid State Chem.*, 2006, **179**, 753–774.
- 66 J. Greedan, N. Raju, A. Wills, C. Morin, S. Shaw and J. Reimers, *Chem. Mater.*, 1998, **10**, 3058–3067.
- 67 J. Wan, P. Ji, B. Li, Y. Xi, X. Gu, L. Huang, M. He and C. Hu, *Chem. Commun.*, 2022, **58**, 589–592.
- 68 Q. Qu, P. Zhang, B. Wang, Y. Chen, S. Tian, Y. Wu and R. Holze, *J. Phys. Chem. C*, 2009, **113**, 14020–14027.
- 69 H. Jiang, P. S. Lee and C. Li, *Energy Environ. Sci.*, 2013, **6**, 41–53.
- 70 T. Brousse, D. Bélanger and J. W. Long, *J. Electrochem. Soc.*, 2015, **162**, A5185–A5189.
- 71 X. Liu, Y. Zheng and X. Wang, *Chem.–Eur. J.*, 2015, **21**, 10408–10415.
- 72 W. Huo, X. Liu and Y. Zhang, *Nanoscale Adv.*, 2019, **2**, 37–54.
- 73 Z. Sun, Y. Zhang, Y. Liu, J. Fu, S. Cheng, P. Cui and E. Xie, *J. Power Sources*, 2019, **436**, 226795.
- 74 M. Huang, F. Li, F. Dong, Y. X. Zhang and L. L. Zhang, *J. Mater. Chem. A*, 2015, **3**, 21380–21423.
- 75 T. Brousse, M. Toupin, R. Dugas, L. Athouël, O. Crosnier and D. Bélanger, *J. Electrochem. Soc.*, 2006, **153**, A2171.
- 76 K. Chen, C. Sun and D. Xue, *Phys. Chem. Chem. Phys.*, 2014, **17**, 732–750.
- 77 K. Naoi and P. Simon, *Electrochem. Soc. Interface*, 2008, **17**, 34.
- 78 J. Shin, J. K. Seo, R. Yaylian, A. Huang and Y. S. Meng, *Int. Mater. Rev.*, 2020, **65**, 356–387.
- 79 B. S. Singu and K. R. Yoon, *J. Alloys Compd.*, 2017, **695**, 771–778.
- 80 N. Tang, X. Tian, C. Yang and Z. Pi, *Mater. Res. Bull.*, 2009, **44**, 2062–2067.
- 81 B. Yin, S. Zhang, Y. Jiao, Y. Liu, F. Qu and X. Wu, *CrystEngComm*, 2014, **16**, 9999–10005.
- 82 M. Xu, L. Kong, W. Zhou and H. Li, *J. Phys. Chem. C*, 2007, **111**, 19141–19147.
- 83 W. Xie, J. Wang, X. Long, X. Wang, S. Zou, L. Zhang, H. Xu, Y. Fu, D. Liu, Y. Li, J. Li and D. He, *J. Energy Storage*, 2021, **35**, 102308.
- 84 L. Athouël, F. Moser, R. Dugas, O. Crosnier, D. Bélanger and T. Brousse, *J. Phys. Chem. C*, 2008, **112**, 7270–7277.
- 85 J. Zhu, W. Shi, N. Xiao, X. Rui, H. Tan, X. Lu, H. H. Hng, J. Ma and Q. Yan, *ACS Appl. Mater. Interfaces*, 2012, **4**, 2769–2774.
- 86 R. E. John, A. Chandran, M. Thomas, J. Jose and K. C. George, *Appl. Surf. Sci.*, 2016, **367**, 43–51.
- 87 M. Khosravi, H. Feizi, R. Bagheri, Z. Song, B. Haghighi and M. M. Najafpour, *New J. Chem.*, 2019, **43**, 4049–4058.
- 88 T. Xiong, Y. Zhang, W. S. V. Lee and J. Xue, *Adv. Energy Mater.*, 2020, **10**, 2001769.
- 89 Y. P. Zhu, C. Xia, Y. Lei, N. Singh, U. Schwingenschlögl and H. N. Alshareef, *Nano Energy*, 2019, **56**, 357–364.
- 90 A. Zhang, R. Gao, L. Hu, X. Zang, R. Yang, S. Wang, S. Yao, Z. Yang, H. Hao and Y. M. Yan, *Chem. Eng. J.*, 2021, **417**, 129186.
- 91 R. Samal, M. Kandasamy, B. Chakraborty and C. S. Rout, *Int. J. Hydrogen Energy*, 2021, **46**, 28028–28042.
- 92 F. Jing, Z. Ma, J. Wang, Y. Fan, X. Qin and G. Shao, *Chem. Eng. J.*, 2022, **435**, 135103.
- 93 D. Rolison and L. Nazar, *MRS Bull.*, 2011, **36**, 486–493.
- 94 H. Wang, J. Zhang, X. Hang, X. Zhang, J. Xie, B. Pan and Y. Xie, *Angew. Chem., Int. Ed.*, 2014, **127**, 1195–1199.
- 95 P. Gao, P. Metz, T. Hey, Y. Gong, D. Liu, D. D. Edwards, J. Y. Howe, R. Huang and S. T. Misture, *Nat. Commun.*, 2017, **8**, 14559.
- 96 L. T. Tseng, Y. Lu, H. M. Fan, Y. Wang, X. Luo, T. Liu, P. Munroe, S. Li and J. Yi, *Sci. Rep.*, 2015, **5**, 9094.
- 97 S. Yao, R. Zhao, S. Wang, Y. Zhou, R. Liu, L. Hu, A. Zhang, R. Yang, X. Liu, Z. Fu, D. Wang, Z. Yang and Y.-M. Yan, *Chem. Eng. J.*, 2022, **429**, 132521.
- 98 X. Zhang, X. Liu, Y. Zeng, Y. Tong and X. Lu, *Small Methods*, 2020, **4**, 1900823.
- 99 H. Z. Chi, Y. Li, Y. Xin and H. Qin, *Chem. Commun.*, 2014, **50**, 13349–13352.
- 100 S. Chen, Y. Xiang, W. Xu and C. Peng, *Inorg. Chem. Front.*, 2019, **6**, 199–208.
- 101 W. Zhang, Y. Yang, R. Xia, Y. Li, J. Zhao, L. Lin, J. Cao, Q. Wang, Y. Liu and H. Guo, *Carbon*, 2020, **162**, 114–123.
- 102 N. Zhang, C. Xu, H. Wang, J. Zhang, Y. Liu and Y. Fang, *J. Mater. Sci.: Mater. Electron.*, 2021, **32**, 1787–1799.
- 103 G. Luo, *Int. J. Electrochem. Sci.*, 2020, **15**, 515–525.
- 104 L. Khandare and S. Terdale, *Appl. Surf. Sci.*, 2017, **418**, 22–29.
- 105 W. Lu, Y. Li, M. Yang, X. Jiang, Y. Zhang and Y. Xing, *ACS Appl. Energy Mater.*, 2020, **3**, 8190–8197.
- 106 L. Kang, C. Huang, J. Zhang, M. Zhang, N. Zhang, Y. He, C. Luo, C. Wang, X. Zhou and X. Wu, *Composites, Part B*, 2019, **178**, 107501.
- 107 M. Huang, Y. Zhang, F. Li, Z. Wang, Alamusi, N. Hu, Z. Wen and Q. Liu, *Sci. Rep.*, 2014, **4**, 4518.
- 108 M. Huang, F. Li, X. L. Zhao, D. Luo, X. Q. You, Y. X. Zhang and G. Li, *Electrochim. Acta*, 2015, **152**, 172–177.
- 109 R. Rajagopal and K.-S. Ryu, *J. Electroanal. Chem.*, 2020, **856**, 113669.
- 110 X. Y. Cao, X. Xing, N. Zhang, H. Gao, M. Y. Zhang, Y. C. Shang and X. T. Zhang, *J. Mater. Chem. A*, 2015, **3**, 3785–3793.
- 111 X. Zheng, Z. Han, W. Yang, F. Qu, B. Liu and X. Wu, *Dalton Trans.*, 2016, **45**, 16850–16858.
- 112 H. Du, C. Wang and J. Lv, *Solid State Commun.*, 2018, **277**, 19–24.
- 113 P. M. Shafi, R. Dhanabal, A. Chithambararaj, S. Velmathi and A. C. Bose, *ACS Sustainable Chem. Eng.*, 2017, **5**, 4757–4770.
- 114 J. Zhu, D. Zhang, Z. Zhu, Q. Wu and J. Li, *Ionics*, 2021, **27**, 3699–3714.
- 115 Z. J. Fan, J. Yan, F. Wei, G. Q. Ning, L. J. Zhi, J. C. Liu, D. X. Cao and G. L. Wang, *ACS Nano*, 2011, **5**, 2787–2794.
- 116 S. Jangu, B. K. Satpathy, M. Raju, C. Jacob and D. Pradhan, *Dalton Trans.*, 2021, **50**, 6878–6888.



- 117 H. W. Chang, Y. C. Huang, J. L. Chen, C. L. Chen, J. M. Chen, D. H. Wei, W. C. Chou, C. L. Dong and Y. C. Tsai, *Catal. Today*, 2022, **388**, 63–69.
- 118 X. L. Bai, Y. L. Gao, Z. Y. Gao, J. Y. Ma, X. L. Tong, H. B. Sun and J. A. Wang, *Appl. Phys. Lett.*, 2020, **117**, 183901.
- 119 J. Xu, K. Hou, Z. Ju, C. Ma, W. Wang, C. Wang, J. Cao and Z. Chen, *J. Electrochem. Soc.*, 2017, **164**, A430–A437.
- 120 H. Lv, X. Gao, Q. Xu, H. Liu, Y. G. Wang and Y. Xia, *ACS Appl. Mater. Interfaces*, 2017, **9**, 40394–40403.
- 121 A. Prasath, M. Athika, E. Duraisamy, A. S. Sharma and P. Elumalai, *ChemistrySelect*, 2018, **3**, 8713–8723.
- 122 W. S. Li, M. L. Chang, K. C. Chuang, Y. S. Li, J. D. Luo and H. C. Cheng, *J. Electrochem. Soc.*, 2019, **166**, A2194–A2198.
- 123 D. Gueon and J. H. Moon, *ACS Sustainable Chem. Eng.*, 2017, **5**, 2445–2453.
- 124 K. Zhao, Z. Xu, Z. He, G. Ye, Q. Gan, Z. Zhou and S. Liu, *J. Mater. Sci.*, 2018, **53**, 13111–13125.
- 125 Y. Chen, C. Kang, L. Ma, L. Fu, G. Li, Q. Hu and Q. Liu, *Chem. Eng. J.*, 2021, **417**, 129243.
- 126 J. Fu, Y. Zhang, H. Zhao, R. Jiang and R. Zhang, *J. Colloid Interface Sci.*, 2020, **559**, 39–44.
- 127 N. Liu, Y. Su, Z. Wang, Z. Wang, J. Xia, Y. Chen, Z. Zhao, Q. Li and F. Geng, *ACS Nano*, 2017, **11**, 7879–7888.
- 128 J. Li, C. Wei, Y. Sun, Q. Liu, X. Zhang and J. Guo, *Adv. Mater. Interfaces*, 2019, **6**, 1801470.
- 129 H. Wang, G. Yan, X. Cao, Y. Liu, Y. Zhong, L. Cui and J. Liu, *J. Colloid Interface Sci.*, 2020, **563**, 394–404.
- 130 M. Cui, S. Tang, Y. Ma, X. Shi, J. A. Syed and X. Meng, *J. Power Sources*, 2018, **396**, 483–490.
- 131 K. Adaikalam, S. Ramesh, P. Santhoshkumar, H. S. Kim, H.-C. Park and H.-S. Kim, *Int. J. Energy Res.*, 2022, **46**, 4494–4505.
- 132 H. Gao, S. Cao and Y. Cao, *Electrochim. Acta*, 2017, **240**, 31–42.
- 133 X. Jia, X. Wu and B. Liu, *Dalton Trans.*, 2018, **47**, 15506–15511.
- 134 M. Jiang, C. Fu, J. Yang, Q. Liu, J. Zhang and B. Sun, *Energy Storage Mater.*, 2019, **18**, 34–42.

