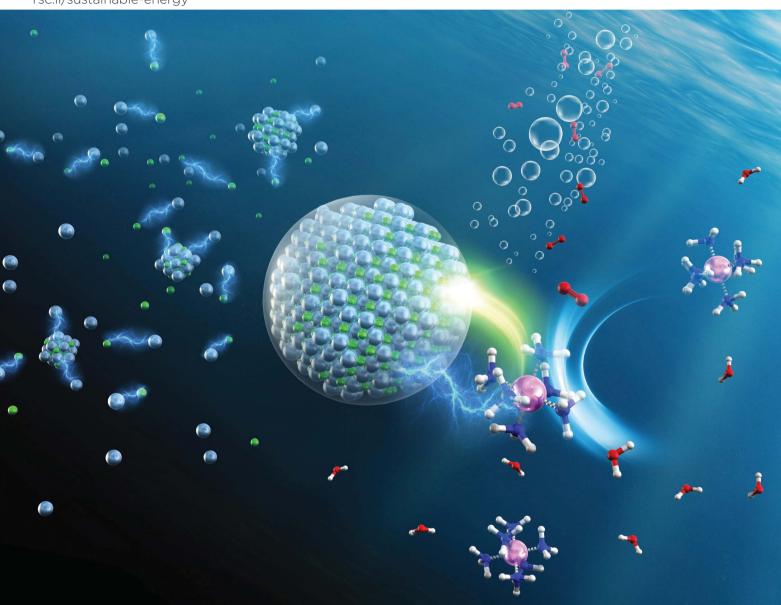
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In situ formation of a molecular cobalt(III)/AgCl photocatalyst for visible-light water oxidation†

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Hexaamminecobalt(III) chloride underwent reaction in an aqueous AgNO₃ solution containing a buffer base (pH 7–9), producing insoluble AgCl that showed no visible light photoresponse. The resultant molecular Co(III)/AgCl system showed water oxidation activity under visible light (λ > 400 nm), approximately five times greater than that of a previously reported CoO_x/anatase-TiO₂ photocatalyst.

Photocatalytic water splitting under irradiation with sunlight is an attractive solution to various energy and environmental issues.1-7 Because O2 evolution requires the transfer of four electrons from two H₂O molecules, it is considered a difficult process from a kinetics perspective, hence necessitating a catalyst. Therefore, the development of a water oxidation catalyst (or photocatalyst) is an important research target in both heterogeneous and homogeneous systems,8-10 even if a sacrificial electron acceptor is employed.11-16

In metal (or metal oxide)/semiconductor composites, one may expect electron transitions from the loaded metal species to the conduction band of the semiconductor, and the metal-tosemiconductor electron transitions can be utilized for water oxidation photocatalysis. 17-19 Our group has reported that modifying a wide-bandgap semiconductor (e.g., TiO₂) with cobalt oxide (CoO_x) or hydroxide (Co(OH)₂) leads to the appearance of a new visible-light absorption band attributable to electron transitions from the loaded Co species to the conduction band of the semiconductor.20-22 For example, Co/ TiO2 hybrid materials are active as photocatalysts for water oxidation in the presence of AgNO3 as an electron acceptor under visible light, whereas unmodified TiO2 is inactive because

In the present work, we report a new visible-light-driven water oxidation system based on a molecular Co(III) complex and a wide-gap semiconductor. We found that hexaamminecobalt(III) chloride ([Co(NH₃)₆]Cl₃) undergoes reaction in an aqueous AgNO₃ solution containing a buffer base (pH 7-9), producing a mixture of insoluble AgCl and aqueous [Co(NH₃)₆]³⁺ species. Surprisingly, the Co(III)/AgCl system, even

Scheme 1 Water oxidation driven by charge transfer from a loaded Co species (e.g., CoO_x) to a semiconductor support. Even though the bandgap photoexcitation of the semiconductor from the valence band (V.B.) to the conduction band (C.B.) under visible light is not possible because of the large bandgap of the semiconductor, interfacial charge transfer from Co to the semiconductor may occur under visible light.

of its large bandgap. We achieved O2 evolution even without physical contact between the Co species and the TiO₂ support. Overall water splitting without a sacrificial reagent was also possible when Co/TiO2 was used as an anode in a photoelectrochemical cell.23 Certain Co compounds are known to be good catalysts (or cocatalysts) for water oxidation.8-12,24,25 Therefore, the Co species in these Co/semiconductor hybrid systems can be regarded as a "catalytic photosensitizer" for water oxidation (Scheme 1), which is unique in heterogeneous photocatalysis. However, the photocatalytic performance of the Co/semiconductor hybrid systems for water oxidation is still unsatisfactory, and the development of more active photocatalysts based on a catalytic Co photosensitizer is strongly desired.

Red Visible light Ox C.B. V.B. CoOx Semiconductor

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without any physical connection, exhibited approximately fivefold greater activity for visible-light water oxidation ($\lambda > 400 \text{ nm}$) compared with the activity of a previously reported CoOx/ anatase-TiO2 photocatalyst.

[Co(NH₃)₆]³⁺ has previously been reported to function as a catalyst for water oxidation in the presence of [Ru(bpy)₃]³⁺ as an oxidant.26 The absorption spectrum of [Co(NH₃)₆]3+ shows a characteristic absorption band in the visible-light region (Fig. 1), which is assigned to d-d transitions (${}^{1}A_{1g} \rightarrow {}^{1}T_{1g}$). AgCl has a cubic rocksalt structure and an indirect bandgap of 3.28 eV.27 The absorption spectrum of our AgCl prepared via a precipitation reaction between AgNO3 and HCl (see the ESI† for details) showed an absorption edge at ~410 nm (Fig. 1 and S1†), with an estimated bandgap of 3.0 eV, which is similar to the reported value. The spectrum of the as-prepared AgCl also showed an increased background in the visible-light region, presumably because of self-sensitization caused by reduced Ag species.28 Although AgCl alone does not efficiently absorb visible light because of its large bandgap, AgCl modified with metals such as Au or Ag has been reported to act as a visiblelight-driven "plasmonic" photocatalyst and/or photoanode.28-30 However, the combination of AgCl with Co species for visiblelight water oxidation has not yet been reported.

The water oxidation reaction was performed at room temperature using a top-irradiation-type reactor connected to a closed gas-circulation system.31 The reaction solutions were prepared by adding different Co sources (e.g., [Co(NH₃)₆]Cl₃, CoCl₂ and Co(NO₃)₂) as "precatalysts" to a 5 mM AgNO₃ solution (140 mL) containing 100 mg of La₂O₃ as a pH buffering agent (pH \sim 8).³² After the reaction solution was degassed, irradiation was carried out using a 300 W Xe lamp and a cutoff filter. The evolved gases were analysed using an online chromatograph.

As shown in Fig. 2a, the [Co(NH₃)₆]Cl₃ and CoCl₂ systems produced O_2 when irradiated with visible light ($\lambda > 400$ nm), with initial rates of approximately 68 and 10 μ mol h⁻¹, respectively. By contrast, the use of $Co(NO_3)_2$ did not yield O_2 . Consequently, among the investigated precatalysts, [Co(NH₃)₆] Cl₃ was determined to be the most appropriate for visible-light water oxidation. Under the optimal conditions, the activity of the [Co(NH₃)₆]Cl₃ system was five times greater than that of the

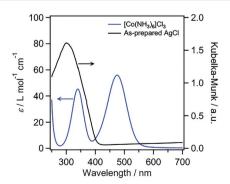


Fig. 1 UV-visible absorption spectrum of $[Co(NH_3)_6]Cl_3$ (5 mM) in H_2O and the diffuse reflectance spectrum of the as-prepared AgCl.

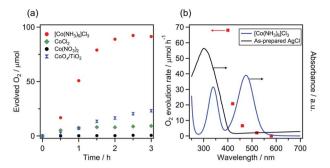


Fig. 2 (a) Time courses of O2 evolution over various precatalyst systems under visible light ($\lambda > 400$ nm) and (b) the dependence of the rate of O₂ evolution by the [Co(NH₃)₆]Cl₃ system on the cutoff wavelength of incident light. Reaction conditions: precatalyst, 30 mg; 5 mM aqueous AqNO₃ solution containing La₂O₃ (100 mg), 140 mL; light source, 300 W Xe lamp with a CM-1 mirror and cutoff filter (L42, Y44, Y48, O54 or R60). In the case of the CoO_x/TiO_2 : catalyst, 100 mg; 10 mM aqueous $AgNO_3$ solution containing La_2O_3 (200 mg). The actual time course data for the O_2 evolution are shown in Fig. S6.†

photocatalyst CoO_x/TiO₂, which was previously reported to exhibit the highest activity among Co/wide-bandgap-semiconductor combinations.22 The initial rate of O2 evolution over the [Co(NH₃)₆]Cl₃ system was comparable to that recorded for representative visible-light-driven photocatalysts (e.g., TaON and LaTiO2N),16 which exhibited high water oxidation activities in aqueous AgNO₃ solution with La₂O₃.

In the absence of La₂O₃, however, the O₂ evolution activity of the [Co(NH₃)₆]Cl₃ system decreased to almost 13% of its value when La2O3 was present (Fig. S2†), which suggests that a pH of \sim 8 is desirable for the efficient production of O_2 . The effect of reaction pH on the O2 evolution activity was investigated in sodium tetraborate aqueous solutions without La2O3. As shown in Fig. S3,† relatively high O2 evolution rates were obtained at pH 7–9. However, O_2 evolution was negligible above pH 10. At pH 10, it was found that the visible light absorption capability of [Co(NH₃)₆]Cl₃ disappeared completely (Fig. S4†), which would be the main reason for the very low activity. These results again indicate the importance of reaction pH. Moreover, no O2 evolution was observed in the absence of either [Co(NH₃)₆]Cl₃, AgNO₃ or light irradiation. The fact that no O₂ evolution was observed in the dark precludes a stoichiometric reaction between the high-valent [Co(NH₃)₆]Cl₃ and H₂O. The apparent quantum yield of the [Co(NH₃)₆]Cl₃ system was 0.20% under irradiation with 460 nm light (Fig. S5†).

Fig. 2b shows the dependence of the rate of O₂ evolution by the [Co(NH₃)₆]Cl₃ system on the wavelength of the incident light. The O2 evolution rate decreased with increasing wavelength of the incident light. O2 evolution was observable under irradiation with >520 nm light but became almost zero under irradiation with >580 nm light. The trend of O2 evolution appears to be associated with light absorption by [Co(NH₃)₆]Cl₃. However, the abrupt decrease observed when comparing the O₂ evolution rate under $\lambda > 400$ nm irradiation with that under $\lambda >$ 420 nm irradiation (Fig. 2b and S6†) strongly suggests that the activity is not simply correlated with the light-absorption

behaviour of $[Co(NH_3)_6]Cl_3$. These results will be discussed again in later sections.

Because the solubility product $(K_{\rm sp})$ of AgCl is 1.8 \times 10⁻¹⁰ mol² L⁻², 33 AgCl should be generated under the present reaction conditions, where 0.8 mM [Co(NH₃)₆]Cl₃ and 5 mM AgNO₃ coexist. The products obtained after photocatalytic reaction for 1 h, at which point the photocatalyst suspension exhibited sufficient activity to generate O₂ (Fig. 2), were characterized by X-ray diffraction (XRD) and UV-visible spectroscopy measurements. As displayed in Fig. 3a, the XRD patterns show that the obtained product was a mixture of La(OH)3, AgCl and Ag, with no peaks assignable to [Co(NH₃)₆]Cl₃. The production of Ag⁰, as a result of the reduction of Ag⁺, was also confirmed (Fig. S7†). Energy-dispersive X-ray spectroscopy analysis was conducted for the solid product; however, Co species were not measurable presumably because of the low concentration. These results strongly suggest that Co species existed predominantly in the solution phase under the investigated reaction conditions, and also preclude the possibility of Co nanoparticle formation. Fig. 3b shows the UV-visible absorption spectra of the reaction solution after filtration of the solid products. The peak intensity at 473 nm was almost unchanged after 1 h of reaction; however, a decrease in the intensity of the absorption at 340 nm (${}^{1}A_{1g} \rightarrow$ ¹T_{2g}) was observed, along with the appearance of a new absorption band at ~300 nm. The 300 nm-band is assigned to the absorption of NO₃ species. ³⁴ A substantial reduction of the

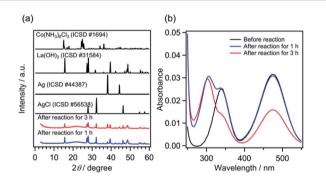
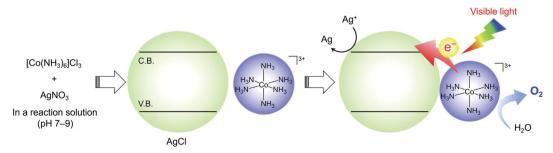


Fig. 3 (a) XRD patterns of the obtained products and (b) UV-visible absorption spectra of reaction solutions after different reaction times. The solution collected after filtration was diluted with H_2O to prepare a 200 mL solution for measurements. The sample "Before reaction" in panel (b) did not include AgNO₃, but only contained 0.56 mM $[Co(NH_3)_6]Cl_3$.

intensity of the visible-light absorption band was observed after 3 h of reaction, indicating the decomposition of $[Co(NH_3)_6]^{3+}$. This decomposition is likely one of the causes of photocatalyst deactivation, but the main cause of the deactivation is the consumption of Ag^+ in the solution (see additional discussion in the ESI†).

On the basis of the aforementioned results, we concluded that the active photosystem consisted of aqueous [Co(NH₃)₆]³⁺ and AgCl as essential components. Note that approximately 0.36 mmol of AgNO₃ still remain even if the formation of AgCl from [Co(NH₃)₆]Cl₃ and AgNO₃ in the initial reaction solution occurs stoichiometrically. Taken together, a possible reaction mechanism for O₂ evolution by the [Co(NH₃)₆]Cl₃ system is proposed (Scheme 2). Under dark conditions before irradiation, AgCl is formed by a reaction between [Co(NH₃)₆]Cl₃ and AgNO₃ in the reaction solution, as confirmed by XRD measurement (Fig. S1†). Visible-light irradiation of the suspension initiates electron transfer from water-soluble [Co(NH₃)₆]³⁺ species to the conduction band of AgCl, followed by the oxidation of water to O₂ and the reduction of Ag⁺ to Ag, which occur on higher-valent (i.e., electron-deficient) Co species and AgCl, respectively. It has been reported that the ζ-potentials of a AgCl single crystal measured in aqueous KNO3 solution (1 and 10 mM) were negative at pH > 5.35 This implies that there is an electrostatic attraction between the in situ formed AgCl and [Co(NH₃)₆]³⁺ species in the reactant solution, which benefits electron transfer from Co to AgCl. It appears, however, that the electrostatic attraction is too weak to bind [Co(NH₃)₆]³⁺ on the AgCl surface. Another possibility that $[Co(NH_3)_6]^{3+}$ species undergo photoexcitation and then inject an electron into AgCl is not realistic because $[Co(NH_3)_6]^{3+}$ is a non-emissive complex (i.e., the lifetime of the photoexcited state is too short to induce a photoinduced electron transfer reaction from its excited state).

The presence of AgCl formed *in situ* in the reaction solution was found to be indispensable for obtaining a measurable amount of O_2 . When AgCl particles were intentionally removed from the solution by filtration prior to photoreaction, the O_2 evolution activity decreased substantially (black squares, Fig. S8†). This confirms that AgCl is essential for achieving visible-light water oxidation with $[Co(NH_3)_6]^{3+}$. When the homemade AgCl was added to the filtered (*i.e.*, AgCl-free) suspension and the photoreaction was conducted, a small amount of O_2 evolution was observed (green diamonds). The addition of NaCl to the filtered suspension produced AgCl



 $\textbf{Scheme 2} \quad \text{Proposed reaction mechanism for visible-light-driven water oxidation over the } [\text{Co(NH}_3)_6]^{3+}/\text{AgCl system}.$

again, resulting in measurable O2 evolution (green squares). However, the thus-obtained activities were much lower than the activity of the standard [Co(NH₃)₆]³⁺/AgCl suspension. As shown in Fig. 2a, the Co(NO₃)₂ system produced negligible O₂. However, the coexistence of NaCl in the Co(NO₃)₂ system led to an improvement of the O_2 evolution rate to 14 μ mol h⁻¹ (blue diamonds, Fig. S8†). In this case, AgCl formation by the reaction between NaCl and AgNO3 was clearly indicated by naked-eye observation of a white precipitate in the solution. All these results indicate the importance of the in situ-generated AgCl, which plays a key role in oxidizing water to O_2 in the presence of Co species. Interestingly, O2 evolution from a mixed system consisting of [Co(NH₃)₆]³⁺ and TiO₂ was very slow, as compared to that recorded for the standard [Co(NH₃)₆]³⁺/AgCl system (Fig. S2†).

Notably, even in the presence of a Cl⁻-containing precatalyst that can form AgCl during the reaction, precatalysts with higher-valent Co (here [Co(NH₃)₆]Cl₃) exhibited better performance (Fig. 2a). This result is likely attributable to the highervalence metal species which are generally advantageous in water oxidation catalysis because of their strong oxidizing ability. Water oxidation in the presence of NaCl without a Co precatalyst hardly occurred as well (orange triangles, Fig. S8†). The AgCl formed by the reaction between NaCl and AgNO₃ showed a small absorption band in the visible region (Fig. S1†). However, it is clear that the visible light absorption of AgCl does not contribute to O₂ evolution. Therefore, if the direct photoexcitation of AgCl occurs at all, it is a minor reaction path during water oxidation by the [Co(NH₃)₆]Cl₃ system.

Conclusions

In summary, we developed a new visible-light-driven water oxidation system that consists of in situ-formed AgCl and $[Co(NH_3)_6]^{3+}$. Although AgCl and $[Co(NH_3)_6]^{3+}$ were individually almost inactive toward visible-light water oxidation, the combination of the two, even without distinct chemical bond linkage, yielded clear photocatalytic activity for water oxidation at pH 7-9 under visible-light irradiation; the activity was five times greater than that of a previously reported CoO_x/anatase TiO₂ system. The results of the present study also suggest that the resultant photocatalytic activities were dependent on the valence state of the Co species initially employed, and that suitable combinations of a molecular catalyst and a semiconductor may give us more active photosystems.

It has been shown that in metal-complex/semiconductor hybrid photocatalysts for CO2 reduction, electron transfer from suspended semiconductor particles to dissolved metal complexes can occur, even without any chemical bond linkage between the two components.36,37 Therefore, the present water oxidation system is another example of the electron transfer type reaction between a molecule and a semiconductor. As mentioned, $\left[\text{Co(NH}_3)_6\right]^{3+}$ has been reported to act as a catalyst for water oxidation,26 but is generally considered to be inactive for ligand exchange. Further investigations of the underlying mechanism and the development of other photocatalytic reactions using this strategy are underway in our laboratory.

Author contributions

The manuscript was written with contributions from all authors. All authors have given approval to the final version of the manuscript.

Conflicts of interest

There are no conflicts to declare.

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